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Chemical Sensing with 2D Materials

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Abstract

During the last decade, two-dimensional materials (2DMs) have attracted a great attention due to their unique chemical and physical properties, which make them appealing platforms for diverse applications in opto-electronic devices, energy generation and storage, and sensing. Among their various extraordinary properties, 2DMs possess high surface-to-volume ratio and ultra-high surface sensitivity to the environment, being key characteristics for applications in chemical sensing. Furthermore, 2DMs' superior electrical and optical properties, combined with excellent mechanical characteristics such as robustness and flexibility, render these materials ideal components for the fabrication of a new generation of high-performance chemical sensors. Depending on the specific device, 2DMs can be tailored to interact with various chemical species at the non-covalent level making them powerful platforms for fabricating devices expressing a high sensitivity towards detection of various analytes including gas, ions and small biomolecules. Here, we will review the most enlightening recent advances in the field of chemical sensors based on atomically-thin 2DMs and we will discuss the opportunities and the challenges towards the realization of novel hybrid materials and sensing devices.

1 Introduction

Since the ground-breaking experiments by Geim and Novoselov on the isolation and study of the outstanding physical properties of graphene in 2004,¹ the research endeavour on two-dimensional materials (2DMs) have grown exponentially becoming a flagship in many classical fields of research such as chemistry and condensed matter physics, and it is particularly blooming in the interdisciplinary realms of nanoscience and material science. One-atom thick graphene is arguably the most glorified material of the last decade; its fascinating physico-chemical properties have spread beyond the academic community and drew attention of world-leading chemical and materials oriented companies as well as public institutions, especially since graphene-based products are going on sale.² This so-called ‘graphene rush’ has triggered the pursuit of atomically thin sheets of other layered materials, such as semiconducting transition metal dichalcogenides (TMDs),³ boron nitride and, more recently, MXenes which include transition metal carbides, nitrides and carbonitrides,⁴ and phosphorene.⁵⁻⁷ The chemical and structural diversity of these 2DMs, whose properties are indeed dictated by their dimensionality,⁸ offer immense opportunities for fundamental and applied research. Different optical and electronic properties may be obtained ranging from the exceptional semi-metallic conductivity of graphene^{9, 10} to the semiconducting characteristics of some TMDs that possess sizeable and tuneable bandgaps, which change from indirect (bulk material) to direct (single layer form).¹¹ This results also in unique photoluminescent properties,¹² thus making such materials suitable for diverse applications such as transistors, photodetectors, electroluminescent devices and luminescence probes. The richness of electronic and optical properties, which can be also engineered by chemical functionalization, combined with their 2D nature, i.e. their extremely high surface-to-volume ratio, make such materials extremely appealing in the field of sensing, ranging from the quantitative detection of gas, metal (alkali and heavy metal), to biologically relevant molecules (e.g. glucose and DNA). Noteworthy, unlike classical digital sensors, 2DMs-based sensors do not possess physical gates for selectively reacting to the targeted species (gas molecules,^{13, 14} metal ions or biomolecules^{15, 16}).

The interaction between 2DMs’ sheets and molecules/ions is accompanied by the adjustment of the properties of both initial components. Such interaction can occur *via* the physical adsorption, i.e. physisorption, of molecular units onto basal plane of 2DMs’ sheets through non-covalent interactions, or through the chemical adsorption, i.e. chemisorption, of reactive

species undergoing chemical reactions with 2DMs to form covalent bonds onto their basal planes. In the field of sensing non-covalent interactions may be preferred when a quick response and a fast recovery rate is required (i.e. real-time monitoring), yet the weakness of the supramolecular forces can be disadvantageous when biomolecules (i.e. enzyme) need to be immobilized on the surface and to be stable during the assay (i.e. in buffered saline solution), thus making covalent linkage more suitable.

The physisorption of molecules onto surface depends on the nature of both analyte and surface; for example, graphene is an extended honeycomb network of sp^2 hybridized carbon atoms characterized by a long-range π -conjugation. Consequently, non-covalent intermolecular interactions involving π -systems are pivotal in the recognition events since subtle changes in the electronic characteristics of the π systems can lead to modifications of the structure and properties,^{17, 18} as well as they may enhance the stability of the physisorbed compounds as observed for proteins, enzyme–drug and DNA–protein complexes.^{19, 20} The understanding of the nature of π -complexes has indeed high importance for graphene based sensors since the gas– π interaction,²¹ H– π interaction,²²⁻²⁵ π – π interaction,²⁶⁻³¹ cation– π interaction,³²⁻³⁹ and anion– π interaction^{26, 40-50} possess different strength which is determined by a combination of attractive and repulsive forces. Compared to graphene, the family of TMDs, which includes molybdenum disulfide (MoS_2), tungsten disulfide (WS_2), molybdenum diselenide ($MoSe_2$) and tungsten diselenide (WSe_2) as the most studied, has not been investigated extensively from this point of view. However, for the case of MoS_2 it has been demonstrated that the physisorption onto its basal plane is mainly driven electromagnetic interactions (e.g. electrostatic and van der Waals).^{51, 52} More generally, 2DMs produced by means of different methods can be very different due to the presence of structural defects which leads to different behaviours, and consequently different performance in the final sensing device. In the last years the term “defects”⁵³⁻⁵⁷ has become a keyword in the field of 2DMs since the presence of defects play a major role in modifying the properties of 2DMs. Although it might have a negative implication when one targets applications in fast opto-electronics, well-designed defects might lead to new and tuneable properties opening a wide range of interesting applications that pristine materials cannot afford such as enhanced electron transfer rate and electrochemical activity as observed for graphene based electrochemical sensors.⁵⁸ Lattice vacancies affect the electronic properties of the TMDs sheets, by lowering their charge carrier mobility and density^{59, 60} as well as triggering photoluminescence^{61, 62} and modifying chemical reactivity.⁶³ In graphene the

defects within the honeycomb network of C=C, whose number depends mainly on the protocol employed for its fabrication, typically consists in point defects, i.e. the absence some sp^2 carbon atoms, and/or the presence of carbon atoms with sp^3 hybridization. The carbon atoms surrounding these defects are electronically perturbed thus they exhibit different electronic structures; therefore, they are chemically activated for further chemical reactions. In other words, the presence of points defects enhances the chemical reactivity of graphene.⁶⁴ sp^2 carbon atoms of graphene can react with highly reactive free radicals, such as those produced by diazonium salts⁶⁵⁻⁶⁷ or benzoyl peroxide;⁶⁸ as well as dienophiles which react with C=C of graphene through a 1,3 dipolar cycloaddition.⁶⁹ However, a massive introduction of defects can be achieved by exposing graphite to strong oxidizing agents⁷⁰ resulting in the formation of graphite oxide (GO). The latter when immersed in water tends to spontaneously exfoliate into single layers of graphene oxide due to the negative charged oxygen functionalities decorating both basal plane and edges.⁷¹ Such material is probably the most investigated 2DM in the field of sensing⁷²⁻⁷⁶ because of its ease functionalization. The negative charges on its surface can interact non-covalently with a wide range of positively charged molecules; for example, outstanding sorption properties for heavy metal cations have been reported.^{77, 78} Moreover, the oxygen functionalities, which consist mostly of hydroxyl and epoxy groups exposed on the basal plane and carboxy and carbonyl functionalities located at the sheet edges,^{79, 80} are extremely reactive thus allowing further modification of GO.^{81, 82} Importantly, the electrical conductivity can be to a great extent restored and tuned during the reduction step of the functionalized graphene oxide (fGO)⁸³ making it an extremely versatile material for sensing with electrical read out. The use of defects to introduce functionalities has been also successfully extended to the family of TMDs where the presence of chalcogen deficiency can be used to decorate the surface of the 2DM with different functionalities. For example reactive sulphur vacancies in MoS₂ may simply arise during the exfoliation process^{59, 84} or be introduced on purpose electrochemically⁸⁵ or by ion-beam irradiation.⁸⁶ These reactive sites can readily react with sulphur-containing moieties such as alkanethiol molecules resulting into a covalent bond formation.⁸⁶⁻⁸⁸ Alternatively, defects can be further expanded into (sub)-nanometre sized pores transforming the 2DMs into permselective membranes⁸⁹ or ultrasensitive sensors even able to sequence DNA.^{90, 91} Such property can be further tuned and harnessed to achieve enhanced permselectivity by controlled chemical functionalization of both pore edges and surface in a post-process chemical treatment.⁸⁹ As already anticipated the 2D nature of such materials offers several advantages in the field of sensing since the atomic thickness provides a direct interaction of

all the atoms with the analyte while the large lateral size not only guarantees a large active surface for the sensing event, but it also facilitates the assembly of the device, as demonstrated for field-effect transistor (FETs), i.e. by enabling a better contact with metal electrodes and better control over the channel structures.⁹² Obviously, the architecture of the sensing device depends on which properties of the 2DMs the analytes is affecting mostly as well as on the nature of both the 2DM and the analyte. Electrochemical sensors based on graphene have been probably the most investigated so far since they provide a direct electrical response. Graphene⁹³ offers indeed a large electrochemical window (up to 2.5 V),⁷² thus enabling the detection of molecules with high reduction or oxidation potential (e.g. nucleic acids), and good electrocatalytic activities for many redox reactions.⁹⁴ Furthermore, due its ambipolar character, the functionalization with both electron withdrawing or donating groups can lead to chemical gating resulting in a change of conductivity of the material.⁹⁵ The combination of the atomic thickness of the 2DMs with the chemical gating that results when the surface potential is changed due to the binding of molecules has led to the generation of new FETs sensors based on 2DMs, as recently reviewed by Mao *et al.*⁹² Interestingly in such a kind of sensors, 2D semiconducting materials outperform their conducting counterparts since the presence of a finite band gap decreases the initial conductance inside the channel thus improving the signal-to-noise ratio. Consequently, the performance of the device is mainly dictated by the band gap which can be tuned by defects engineering, doping, as well as by playing with the thickness of the material opening a wide range of opportunities as recently demonstrated by Cui *et al.*⁹⁶ in a phosphorene-based FET gas sensor.

While a direct electrical response is generally preferred for practical applications, the interactions of the 2DMs with analytes give rise to interesting optical phenomena such as the modulation of the photoluminescent properties opening up a wide range of opportunities. Graphene and GO are known to be highly efficient fluorescence quenchers if compared to organic compounds, thus Förster resonance energy transfer (FRET) sensors have attracted increasing interest in the last years especially for biomedical applications since they can be used to measure precisely nanometre-scale distance and changes both *in vivo* and *in vitro*⁹⁷ resulting into nanobiosensors with excellent sensitivity, selectivity, and biostability.⁹⁸ Even single-layer MoS₂ nanosheet possesses high fluorescence quenching efficiency and by taking advantage of such a characteristic it has been exploited as sensing platforms for the detection of DNA and small molecules.⁹⁹ However, the use of 2DMs in FRET is not limited to energy acceptor since a proper functionalization may result in photoluminescent flakes which can act

as energy donor and be quenched by more electron deficient molecules such as nitro compounds which are common constituent to prepare powerful explosives.¹⁰⁰

The use of 2DMs as substrate for enhancing the Raman signals of adsorbed molecules represented a major breakthrough in the field of sensing.^{101, 102} Different Raman vibrational modes can be enhanced depending on which layered material the molecule is adsorbed onto. Surface Enhanced Raman Spectroscopy (SERS) is mainly employed to explore the detection of chemical and biological species¹⁰³⁻¹⁰⁶ due its high sensitivity (even down to single molecules)^{107, 108} and the bar-code like reading that comes from the narrow vibrational bands in the Raman spectrum. Different Raman enhancement mechanisms have been proposed for different 2DMs, however, like in the previous examples, surface modification¹⁰⁹ as well as the thickness¹¹⁰ play a fundamental role in terms of selectivity and Raman signal enhancement.

The intensive research on 2DMs for sensing application has been further motivated by their intrinsic mechanical properties such as robustness, flexibility, lightweight which makes possible the realization of portable and wearable sensors with tremendous impact on our society enabling the monitoring of the wearers' health, fitness, and their surroundings.¹¹¹ The development of wearable chemical sensors faces multiple challenges on various fronts such as power, analytical procedure, communication, data acquisition, processing and security. Nevertheless, several examples of flexible graphene-based wearable gas and chemical sensors has been recently reviewed¹¹² and also a wearable patch for sweat-based diabetes monitoring and feedback therapy has been reported,¹¹³ which combines a heater, temperature, humidity, glucose and pH detection.

Looking forward to the emergence of portable and ultrasensitive sensors based on 2DMs, in this Review article we discuss the recent advances in gas, alkali and heavy metals sensing as well as relevant chemical entities, emphasizing the performance of the different sensor devices based on 2DMs in terms of sensitivity, selectivity, robustness and response time, by focusing on the device preparation and their suitability as wearable sensors. Each section will start with a general introduction to the sensing towards a specific analyte (gas, metals or chemically relevant molecules), and will follow with a detailed discussion on the smartest approached strategies and the best recent achievements obtained for various 2DMs ranging from graphene to MXenes. For each 2DM we classify the sensors on the basis of the type of signal transduction; in particular, we focus on electrochemical, FET, fluorescence and SERS

sensors. While presenting remarkable examples of chemical sensors we will provide specific attention to the most important figures of merit such as the sensitivity, selectivity, robustness, response time, and to the use of strategies to minimize and ideally exclude effects of interfering analytes, with the final aim of developing flexible sensors for wearable technology.

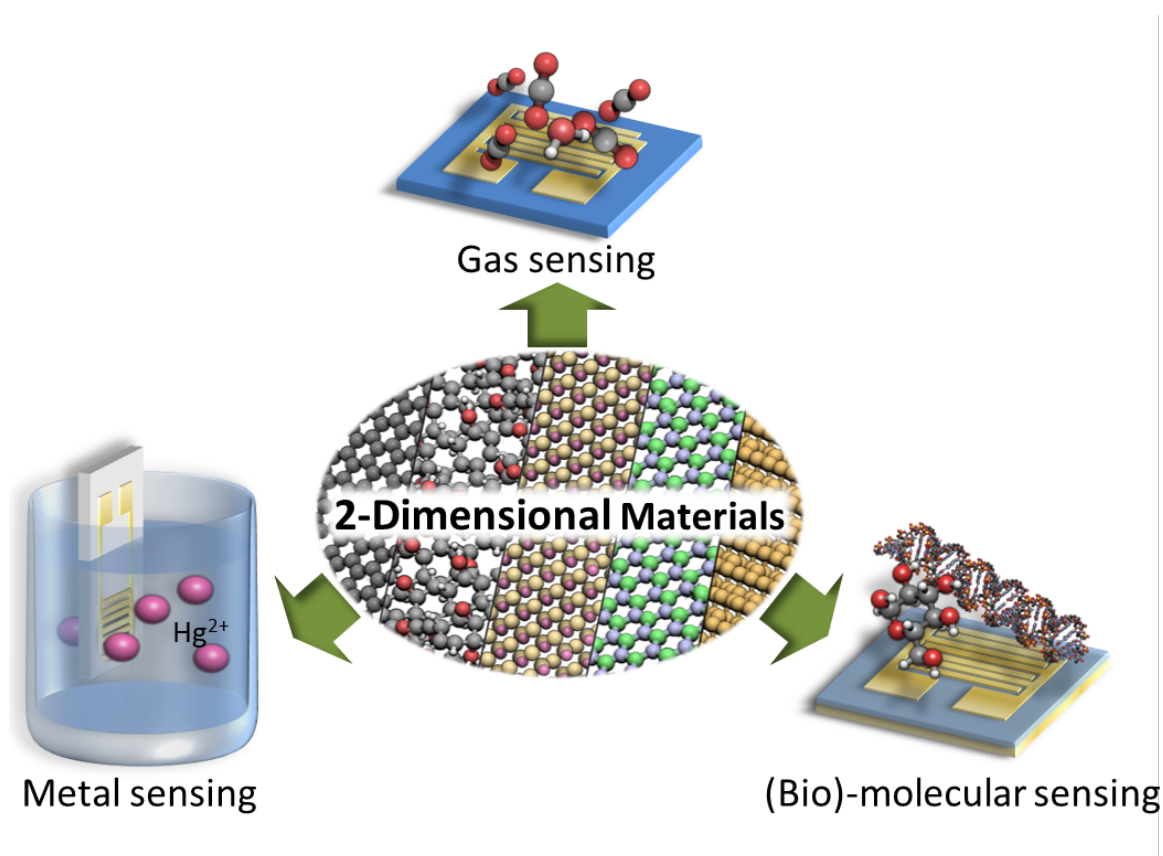


Figure 1: Schematic representation of the different molecular chemical sensors based on 2DMs approaches that have been explored over the last years.

1.1 Overview on the properties of 2DMs

Graphene is an atomically thin, planar membrane of carbon atoms arranged in a honeycomb lattice whose unique properties were firstly investigated by Geim and Novoselov in 2004.¹ Graphene can be seen as a single layer of graphite, and while the latter behaves metallically, graphene is a semi-metal featuring a unique zero band gap.¹¹⁴ In addition graphene exhibits remarkable thermal and electrical conductivity and an impressive mechanical strength being

superior to steel.² Graphene is particularly promising for sensing by virtue of its extremely high conductivity, and its large surface area. Despite being only one atom thick graphene is impermeable in its pristine form, effectively blocking the passage of even the smallest molecules.¹¹⁵ The impermeable nature of graphene has triggered extensive studies on its application as a barrier for liquid and gas permeation,^{116, 117} as well as on its use as shielding material protecting metallic surfaces against corrosion.¹¹⁸ Such unique property of graphene has triggered extensive efforts towards the use graphene and other 2DMs for the design of ultrathin water-separation membranes and as platforms to absorb (heavy) metal ions. Hitherto the highest quality graphene flakes have been obtained by using a *top-down* approach relying on the mechanical exfoliation. The graphene layers are most commonly exfoliated from the bulk graphite *via* the “scotch tape” method.¹¹⁹ Unfortunately, the flakes produced with such a methods have high quality, yet very limited lateral size. Large area and high-quality mono- and few-layer thick graphene sheets can be obtained by two *bottom-up* approaches: the chemical vapour deposition (CVD) and the epitaxial growth. CVD graphene is usually obtained through a catalytical decomposition of hydrocarbons (usually methane) on a hot (~1000 °C) metal surface (Cu, Ni and Co) under vacuum.¹²⁰ Since the graphene grows directly onto the metal surface, different techniques have been developed to transfer it onto dielectric substrates.¹²¹ Epitaxial growth is another method to obtain large and uniform high-quality graphene films.¹²² Typically, SiC is heated under high vacuum at high temperature (> 1200 °C); this allows the surface silicon atoms to evaporate, yielding the rearrangement of the carbon atoms to form a graphene layer. Mechanical exfoliation makes it possible to obtain high quality graphene monolayers, but the very low throughput and yield of the so-obtained graphene flakes hinder any industrial application, while CVD and epitaxial grow allow to obtain large-area graphene monolayers, although the production costs remain high. The production of GO followed by its reduction is a high throughput, easily scalable and cheap method to obtain large amount of graphene.¹²³ It is however fair to note that reduced graphene oxide (rGO) is less conductive and have numerous structural defects and residual functional groups compared to pristine graphene, yet their presence could offer a clear route for improvement of the sensing capabilities as already discussed.¹²⁴ Similarly GO has a higher number of oxygen functional groups, but it is an electrical insulator.¹²³ A compromise in terms of quality of the flakes and yield is offered by the liquid-phase exfoliation (LPE) of graphite: graphene flakes dispersed in water/surfactants or organic solvents can be obtained with the aid of ultrasonication,^{125, 126} shear mixing¹²⁷ or electrochemical approach.¹²⁸

Alongside graphene, transition metal dichalcogenides are the most studied 2DMs. TMDs are semiconductors of the type MX_2 , where M is a transition metal atom (such as Mo, W, Re and others belonging to the 4, 5 and 10 group) and X is a chalcogen atom (such as S, Se or Te); these materials that exist in the bulk form can be exfoliated yielding 2D monolayers in which one atomic layer of the metal is sandwiched between two layers of X atoms. TMDs are promising materials for the use in electronic devices,^{129, 130} energy storage¹³¹ and sensors¹³²⁻¹³⁴ due to their unique chemical and physical properties, including semiconducting properties, high surface-to-volume ratios and absorption coefficients, adjustable and direct band gaps and availability of reactive sites for redox reactions.¹³⁵ The chemical versatility of TMDs and their reactivity together with their natural abundance result in the ever growing interest in those materials.¹³⁵ Although majority of TMDs exhibit similar structure, the opto-electronic properties are diverse. In particular, the band gap of TMDs ranges from the insulating HfS_2 to the conductive VSe_2 .¹³⁵ Noteworthy, 2D TMDs exhibits two different crystal phases: a trigonal prismatic and a octahedral phase, usually denoted as 2H and 1T phases, respectively. In Mo- and W-based TMDs the 2H phase is thermodynamically stable and has semiconductive properties, while the 1T phase displays metallic properties. Other TMDs present a stable distorted octahedral phase, usually referred as 1T' phase.¹³⁶ The most investigated TMDs are MoS_2 , MoSe_2 , WS_2 , WSe_2 , ReS_2 and ReSe_2 . Among these, MoS_2 is considered as the most promising material as it possesses various physico-chemical properties including fast electron transfer, good conductivity and quenching ability and it is abundant in nature as molybdenite. Multiple preparation techniques were reported to obtain nanosized MoS_2 including LPE,¹³⁷ chemical vapour deposition,^{138, 139} lithium intercalation.^{99, 140, 141} LPE makes it possible to obtain dispersion of MoS_2 nanosheets and other TMDs in higher yield and low costs. However, the obtained materials are usually quite thick and feature limited lateral dimensions.¹⁴² Larger mono and few layers thick MoS_2 of good quality can be produced by CVD¹⁴³ directly on various substrates, including Si/SiO₂.¹⁴⁴⁻¹⁴⁷ The semiconducting MoS_2 exhibits a thickness-dependent band gap, ranging from 1.3 eV for bulk MoS_2 to 1.9 eV for isolated monolayers,¹⁴⁸ good conductivity and fast electron transfer. MoS_2 is currently considered as one of the most preferable materials for chemical sensing applications, predominantly due to available edges that facilitate electron transfer.^{3, 149} Moreover, it can be easily produced in large scale and dispersed in numerous solvents to obtain desired structures.^{137, 150} Layered tungsten disulphide (WS_2) nanosheets are one among the newly emerging TMDs, which consist of S-W-S sandwich structures. WS_2 nanosheets have been the focus of intense research effort as catalyst, semiconducting material for field-

effect transistors, active material in lithium-ion battery, etc. The ambipolar WSe₂, has a similar trigonal lattice and a 1.6 eV band gap (in the monolayer). Finally ReS₂, ReSe₂ and PtSe₂ are recently explored TMDs, characterized by a highly distorted triclinic structure which confers to them extremely anisotropic optical, mechanical and electrical properties.¹⁵¹ Mono and few-layer ReS₂ and ReSe₂ can be obtained by mechanical exfoliation and CVD.¹⁵²

Similar in structure to TMDs, SnS₂ is a layered material possessing a triclinic structure which can be exfoliated in mono and few-layers with semiconducting properties.¹³⁵ Compared to TMDs, SnS₂ exhibits a higher electronegativity, which can potentially enhance the absorption sites and the sensing capabilities.

Among the emerging 2DMs black phosphorus (BP) has been recently re-discovered as a 2DM under the commonly used name of phosphorene. This material, which is, in analogy of graphene, a single honeycomb layer of the layered phosphorus allotrope can be obtained *via* mechanical or liquid exfoliation and presents remarkable electrical properties of tuneable direct band gap and high mobility.¹⁵³ BP exhibits relatively low energy band gap, which depends on the number of layers and ranges from 0.3-1.5 eV.^{154, 155} BP exhibits fascinating optical and electronic properties such as size dependent optical response or anisotropic electrical conductivity, which simplifies characterization and segregation methods and makes BP promising material for various electrical applications.^{156, 157} BP has already found several applications as FETs, photodetectors, solar cells and gas sensors.¹⁵⁸⁻¹⁶¹ However, its application in devices is limited by the very low material's stability when exposed to air or water.¹⁶² Such a negative characteristic can become an advantage when fabricating chemical sensors.

Hexagonal boron nitride (hBN) is a synthetic polymorph of boron nitride with layered structure analogue to graphite. Similarly to graphene, hBN WS₂ exhibits a two-dimensional honeycomb-like structure with strong covalent bonds in the plane. It is characterized by relatively high band gap ranging from 3.6 to 7.1 eV, which determines its insulating properties.^{163, 164} This feature provides transparency in the visible and near-IR regions extending its application potential.¹⁶⁵ The weakly bonded layers can be exfoliated down to monolayer hBN, which has electrical insulating properties, but also high thermal conductivity, mechanical strength, hardness and chemical stability.¹⁶⁶ The production method strictly determines the 2D hBN nanosheets crystallinity, structure and other physical and chemical properties. Multiple methods are exploited for its production including mechanical

exfoliation,^{167, 168} liquid exfoliation^{137, 169-171} and chemical vapour deposition.¹⁷²⁻¹⁷⁴ Features such as good thermal conductivity, low toxicity, high mechanical strength, atomically smooth surface makes BN good candidate for sensing applications.

The large family of 2DMs also includes early transition metal carbides/nitrides (MXenes). This novel class of 2DMs shows promising properties toward a range of applications. MXene nanosheets originates from the MAX phases which constitute a class of nitrides and carbides in which M is an early transition metal, A is an element from group 13 or 14 of the periodic table, X is either C and/or N. The MXene layers are obtained from the MAX phase by the removal of the element A, which is intercalated between MXenes layers of the crystal, by treatment with HF and further sonication.¹⁷⁵ MXenes display good electronic conductivity and unique morphology making them suitable for energy storage,¹⁷⁶ electrochemical capacitors¹⁷⁷ and chemical sensors.^{178, 179}

In the field of sensing it is generally preferred to make use of non-covalent interactions between the sorbents, i.e. the active material, and the analyte to ensure a quick response and a fast recovery rate (i.e. real-time monitoring). The use of pristine 2DM nanosheets has some drawbacks. A major limitation is represented by the poor porosity, which limits the number available recognition sites. In particular, 2DM nanosheets produced *via* LPE exhibit a strong tendency to self-aggregate and form even micron-sized stacked structures. Moreover, nanosheets produced *via* LPE exhibit low charge carrier as well as low field-effect mobility.

Alongside the capacity of pristine 2DMs to detect various analytes, extensive efforts have been devoted towards the development of 2DMs-based sensors through the non-covalent functionalization of the 2DMs with both inorganic and organic moieties which act as spacers/pillars imposing a certain distance between adjacent sheets. This results in an enhanced porosity of 2DM-based composite which determines a greater sensitivity for the analyte of choice. Such an approach enables the tailoring of the properties of 2DM-based sensors, which could preserve many of the unique characteristics of the individual 2DM sheets and benefit from the presence of (in)organic moieties. Such moieties, besides acting as separators, can by design incorporate the receptor of the analyte of choice, endowing highest selectivity in the recognition and sensing process. In this context, the fabrication of 2DM-based sensors through non-covalent interactions between individual 2DM sheets is extremely appealing, as it could result in the structures exhibiting a remarkable enhancement of sensitivity towards specific analytes.

2 Applications in gas sensing

A gas sensor is a device that can detect the presence and quantify the concentration amount of a specific gas in the atmosphere such as water vapour (humidity), organic vapours and hazardous gases.¹⁸⁰ Gas sensors have attracted a strong interest and are widely employed in environmental monitoring and emission control,¹⁸¹ personal and military safety,¹⁸² production control in agriculture and industry and medical diagnostics.¹⁸³⁻¹⁸⁵ Among variety of materials used for gas sensing devices, metal oxides have been largely considered for their high sensitivity and low costs; however, their crucial drawbacks as high temperature operation and large energy consumption, as well as poor selectivity, have hindered their practical application.¹⁸⁶ Therefore, many efforts have been devoted in developing new gas sensors with high sensitivities and low operating temperatures. Numerous active materials have been considered including conducting polymers¹⁸⁷ and carbon nanotubes (CNTs),¹⁸⁸ which possess interesting low operating temperatures but suffer from long response and recovery time, poor stability, degradation and difficult processing.

In the last few years 2DMs are emerging as new materials for gas sensing by virtue of their unique properties, which promise to largely improve the sensitivity of the sensors. Among 2DMs, graphene is particularly appealing, because of its low resistivity and electrical noise.¹⁸⁹ Such characteristics enable the detection of small changes in the intrinsic resistance resulting from the interaction of graphene with gaseous species. Alongside graphene, also its analogues such as TMDs (MoS_2 ,¹⁹⁰⁻¹⁹³ WS_2 ,^{194, 195} MoSe_2 etc.), layered metal oxides (MoO_3 ,¹⁹⁶ SnO_2 ^{197, 198}), phosphorene,¹⁵⁸ and h-BN¹⁹⁹ etc. have also been employed for gas sensing by exploiting their thickness dependent characteristics and semiconducting properties. Indeed, the presence of a band gap that can be modulated by the interaction with gases is a tool that enables enhanced sensitivity. Furthermore, the semiconducting properties of these materials render them suitable for being integrated as active components in FETs, permitting low power consumption and miniaturized devices.¹⁹⁹ In these devices, the adsorption of gases on the semiconducting material determines a change in their conductivity, which can be measured as a variation in the drain current.¹⁴⁴ Other common gas sensing devices are represented by chemiresistors¹⁸⁵ and chemi-capacitors,²⁰⁰ in which the sensing material is interposed between two electrodes, and the gas molecules adsorbed on the surface of the material induce a change in the resistance or capacitance, which can be directly quantified.

The performances of the gas sensors are usually characterized by various figures of merit, with the most important being sensitivity, selectivity, stability, response, and recovery time as well as the costs, dimensions and flexibility. In this chapter we will define the response $R(\%)$ of a given device as the ratio between the difference in resistance (or another output) when it is operating in the presence or in the absence of the sensed gas:

$$R(\%) = \frac{R(i)-R_0}{R_0} * 100 \quad (\text{eq. 1})$$

In this part of review we will provide an overview over the gas sensing mechanism for devices containing active 2DMs, and we will present the most interesting and up-to-date gas sensing applications of graphene, its derivatives, TMDs and other 2DMs.

2.1 Graphene-based gas sensing

2.1.1 Gas sensing with pristine graphene

Graphene sheet exhibits two exposed surfaces without a bulk separation, thus it has all its atoms exposed to the environment; such a structure makes it an ideal building block for the detection of gas molecules.²⁰¹ Furthermore, the high variability of its resistance and low electrical noise of some devices makes it possible to detect even extremely small changes in the number of carriers generated by the interaction with gas molecules.¹²⁴ The most frequent detection mechanism of graphene, as well as many other 2DMs, is based on charge transfer process, in which the sensing material and the gas molecules behave as charge donors or acceptors.²⁰² In particular, Leenaerts *et al.*²⁰³ used density functional theory (DFT) calculations to demonstrate that gas molecules adsorbed on the surface of graphene act as a dopant changing the carrier density of graphene. The interaction energies of H₂O, CO, NO, NO₂ and NH₃ with graphene were calculated, showing that H₂O and NO₂ molecules behave as strong electron acceptors for graphene, thereby resulting in a decrease of charge carriers in the latter, while NH₃, NO and CO behave as electron donor enhancing the concentration of charge carriers. Upon exposure of the graphene surface to air or to an inert gas the sensed gas molecules are promptly desorbed and the initial conductivity of graphene is recovered.

The first application of graphene as a gas sensor was reported in 2007 by Novoselov *et al.*²⁰¹ A micromechanically exfoliated graphene flake supported on Si/SiO₂ substrate was firstly etched in a bar shape and patterned with top gold contacts exploiting Electron-Beam

lithography (EBL) (Figure 2a) and later exposed to H₂O, NO₂, NH₃ and CO gases. It was shown that the presence of gas molecules causes a doping in the graphene changing the concentration of charge carriers. In particular, it was demonstrated that H₂O and NO₂ decrease the charge carrier concentration while an increase of carrier concentration was observed when NH₃ or CO were employed (Figure 2b). Furthermore, exploiting the Hall resistivity, changes in the number of carriers of just a single electron were investigated in the vicinity of the Dirac point. In this way, by exposing the device to very diluted NO₂ gas it was possible to sense the adsorption of a single NO₂ molecule. In fact, it can be seen that the Hall resistivity varies in a step-like manner, because each NO₂ molecule adsorbed/desorbed produces a variation of one electron in the number of carriers (Figure 2c). This outstanding performance was attributed to the high carrier mobility and extremely low noise of the so patterned graphene device.

In 2009 Dan *et al.*²⁰⁴ demonstrated that the outstanding sensing performances of Novoselov's device could not only be ascribed by the use of pristine graphene, whose inert surface is not optimal for adsorbing molecules, but it can also be attributed to the lithographic process employed for device fabrication. Towards this end, a miniaturized FET based on mechanically exfoliated graphene on Si/SiO₂ substrate was fabricated by EBL and used to detect various gas vapours, including H₂O, nonanal, octanoic acid, and trimethylamine down to the ppm level. It was claimed that the sensitivity of the device was enhanced by the presence of a contamination layer on the graphene surface induced by EBL process. This layer consists of residues such as the photoresist, acting as an absorbent layer that concentrates analyte molecules onto the graphene surface. In fact once this layer was removed by Ar/H₂ etching, the device became almost completely insensitive to the analysed vapours. A similar device based on mechanically exfoliated graphene patterned through EBL was developed by Ko *et al.* to detect NO₂, with a response of 9 % to 100 ppm of NO₂.⁸⁰

EBL was also employed by Balandin *et al.*²⁰⁵ to fabricate a FET that can selectively detect different organic vapours through analysis of the low frequency noise of the device. In particular, it was shown that the gas molecules act as trap centres, leading to fluctuation in the carrier concentration and thus noise. In addition, the different kinetics of adsorption/desorption of the different vapours lead to different characteristic noises and made it possible to selectively distinguish the presence of tetrahydrofuran (THF), methanol (MeOH), acetonitrile (MeCN) and chloroform in the environment.

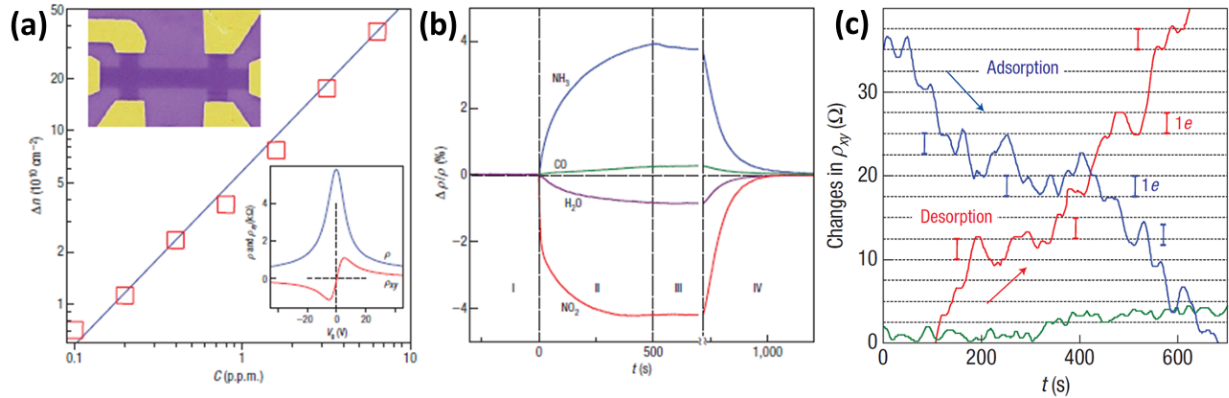


Figure 2. (a) Dependence of the concentration of charge carriers (Δn) in single-layer graphene when exposed to different concentrations (C) of NO_2 . Upper inset: Scanning electron microscope (SEM) image in false colour of this device (the device is $1 \mu\text{m}$ wide). Lower inset: Characterization of the graphene device by using the electrical field-effect. (b) Changes in resistivity caused by graphene's exposure to 1 ppm of various gases highlighting the positive or negative doping effects. (c) Changes in Hall resistivity observed near the Dirac point during adsorption (blue curve) and desorption (red curve) of strongly diluted NO_2 . The green curve is a reference of the same device exposed to pure He. The grid lines correspond to changes in ρ_{xy} caused by adding one electron charge. Figure adapted from Ref. ²⁰¹ with permission from Springer Nature.

Graphene obtained by mechanical exfoliation was employed also by Yoon *et al.* ²⁰⁶ as a sensor of CO_2 . In order to minimize the number of residues resulting from the traditional “scotch tape” cleavage method and from the EBL process, a monolayer of graphene was exfoliated from the bulk by employing a polydimethylsiloxane (PDMS) stamp and transferred on top of gold electrodes pre-patterned on a Si/SiO₂. Such a device showed moderately fast (~ 10 s), reversible, high and linear response to a concentration of CO_2 ranging between 10 and 100 ppm at room temperature. A similar device based on mechanically exfoliated graphene was developed by Ko *et al.* to detect NO_2 , with a response of 9 % to 100 ppm of NO_2 .⁸⁰

Larger and high-quality graphene sheets can be obtained by chemical vapour deposition:¹²¹ such graphene type has been used to detect NO_2 ,²⁰⁷⁻²¹⁰ NH_3 ,²¹⁰⁻²¹² CO_2 and humidity.²¹³ CVD graphene devices showed in fact good sensitivity towards the detected gases, yet it suffered from difficult desorption of the sensed gases and exhibited slow and incomplete recovery at low temperatures. CVD graphene monolayers transferred onto Si/SiO₂ substrates were used

by Gautam *et al.* to sense NH_3 , CH_4 and H_2 .²¹¹ The best performance of the device in terms of sensitivity was observed when operating at 150-200 °C, yet the use of such extreme experimental conditions was not accompanied by an improvement in the response time. Choi *et al.*²⁰⁹ realized a transparent and flexible NO_2 detector, in which single layer graphene (SLG) was employed as a sensing active component and bi-layer graphene (BLG) was used as an internal heater: by applying an electrical power of 1.7 W to the device it was possible to heat the device itself up to 200 °C for joule effect. Such a device showed a response of 39 % to 40 ppm of NO_2 and the presence of the heaters was found to accelerate the recovery time of the device reaching values of a few seconds. Kim *et al.*²¹⁰ in 2015 presented a similar self-heated, transparent, flexible NO_2 sensor. CVD SLG was patterned on the Cu foil to obtain a conductive channel of 5 μm width between two graphene electrodes and then transferred onto a flexible Polyimide (PI) substrate (Figure 3a). Also in this case, the SLG micro channel was self-heated up to 73 °C by applying a bias of 60 V, yielding an increase in the sensitivity and recovery time of the device ($\Delta R/R_0 = 12\%$ and recovery time of 82 s to 5 ppm of NO_2 at 60 V) (Figure 3b-c). This device exhibited also good selectivity, reversibility and durability even under mechanical bending and negligible influence of humidity. Yavari *et al.*²¹² in 2011 realized a sensor based on unconventional CVD graphene. In particular, graphene foam was produced by growing CVD graphene onto a porous nickel template, which was dissolved afterwards to obtain the self-standing foam. After contacting the foam with two electrodes a 30 % response to 1000 ppm of NH_3 was measured, yet the response and the recovery time were both found being slow (~ 1000 s).

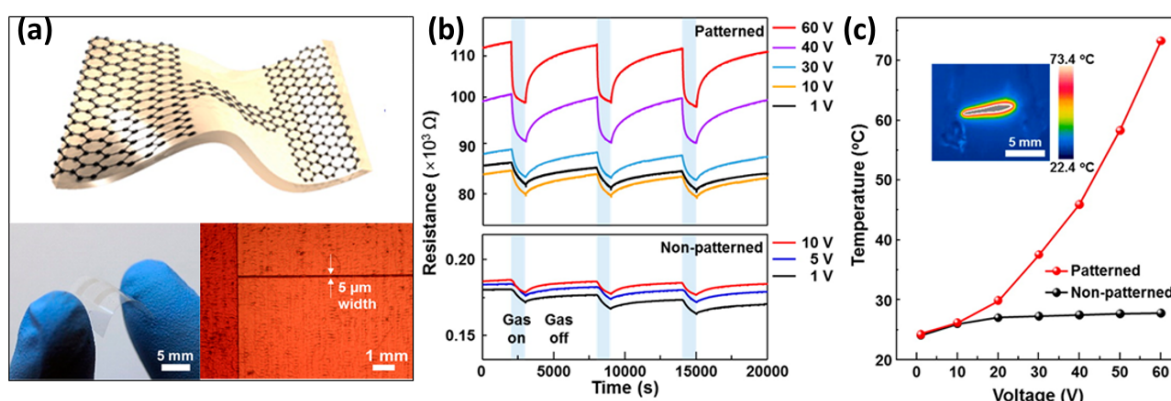


Figure 3. (a) Scheme (top) and photograph (bottom left) of the all-graphene flexible, transparent gas sensor and micrograph of patterned graphene (bottom right). (b) Response curves, i.e. resistance changes, recorded in patterned and non-patterned graphene sensors when exposed to three pulses of 5 ppm of NO_2 . (c) Thermographic image and thermal

characteristics at different bias voltages. (Figure adapted from Ref.²¹⁰ with permission from the American Chemical Society)

The sensing performance can be tailored by the choice of the substrate in epitaxially grown graphene.²¹⁴ In fact SiC possesses a wide number of different polymorphs and each one of them, as well as diverse crystal faces, can exhibit doping effects on graphene.²¹⁵ As an example Pearce *et al.*²¹⁶ realised a NO₂ sensor based on epitaxially grown SLG on the atomically flat Si face of 4H-SiC. This sensor displayed an extremely high response to 2.5 ppm of NO₂ ($\Delta R/R_0 \approx 120\%$). Interestingly, the response switched from *n*-type to *p*-type at increasing NO₂ concentration because of the intrinsic *n*-doped nature of graphene grown on 4H-SiC. This behaviour was also demonstrated by Nomani *et al.*²¹⁴ In their work graphene grown on C face of 6H-SiC and on Si face had opposite behaviour when exposed to 18 ppm of NO₂: the one grown on C face was *n*-type and exhibited a 4.5 % increase of the conductance, while the other was *p*-type and exhibited a 10 % decrease of the conductance.

In order to enhance the performance of graphene-based gas sensor a promising solution consists in introducing defects and functional groups. In fact, the absence of dangling groups on graphene's surface renders it inert to the chemisorption of gas molecules. For example, the ozone treatment of graphene sheets is a viable approach to introduce oxygen functional groups uniformly on the surface. Such defects can act as recognition sites and promote the interaction with gaseous molecules through hydrogen bonds and van der Waals forces.²¹⁷ Chung *et al.*²⁰⁸ treated CVD SLG with ozone and demonstrated that the so-treated graphene features better performances in terms of sensitivity and response time for detecting NO₂ compared to pristine CVD graphene. Along the same line, Masel *et al.*²¹⁸ proved that the CVD polycrystalline graphene was much more sensitive to various organic vapours than monocrystalline pristine graphene because of the presence of structural defects.

2.1.2 Gas sensing with graphene oxide and reduced graphene oxide

The key role played by oxygen containing functional groups in graphene oxide is perfectly highlighted by its sensing capabilities. The sensing performances of GO can be easily tailored by varying the amount of its oxygen containing functional groups *via* reduction (thermal or chemical) towards rGO. These unique characteristics make pristine and reduced GO the most studied platforms for gas sensing.¹²⁴ Numerous articles on GO,^{200, 219} rGO^{220, 221} and

functionalized GO (fGO)^{222, 223} based gas sensors have been published. The first application of rGO as gas sensor was developed by Robinson *et al.* in 2008.⁷³ GO was spin-coated on the top of a Si/SiO₂ substrate and Ti/Au pads were evaporated as top electrodes (Figure 4a). The device was exposed to hydrazine vapours at 100 °C to trigger the reduction of GO to rGO. Figure 4b portrays the response of the device when exposed to pulses of acetone vapours for 5s. The response (e.g. electrical) of the device varies depending on the duration of the reduction treatment. In particular, it is possible to observe two response components: a fast response, which increases with the reduction time due to the decrease of the number of oxygen functional groups, and a slow response, which decreases with increasing reduction time. It was explained that the fast response is the result of the gas adsorption onto low-energy binding sites, such as sp²-bonded carbon, while the slow response occurs from the interaction of gases with higher-energy binding sites, such as structural defects, and oxygen bearing functional groups. Furthermore, the electrical noise of the device can be pushed to the limit of the thermal noise (Figure 4c) for very thin films and long reduction treatment, being much lower of CNTs network. Thanks to the low electrical noise the limit of detection of various vapours was found being extremely low. For example, it was possible to sense 70 ppb of HCN and 5 ppb of dimethyl methylphosphonate (DMMP).

A flexible all printed gas sensor based on rGO reduced with ascorbic acid was realized by Manohar *et al.*²²⁴ This was done by preparing a dispersion of rGO in water in presence of a surfactant, followed by its ink-jet printing onto a PET substrate. The device displayed a response with a detection limit on the tens of ppm range; such a response is positive, i.e. displaying increasing resistances, for NH₃, dichloromethane and various alcohols, whereas it was found providing negative response for Cl₂ and NO₂. Such phenomenon has been attributed to the electron withdrawing character of Cl₂ and NO₂, which enhance the conductivity of semiconducting *p*-type rGO.²¹² In another case,²²⁵ a NO₂ gas sensor was fabricated by reducing and patterning *via* laser irradiation the GO on various flexible substrates. The device exhibited a good sensitivity, which could allow the detection of 20 ppm of NO₂, with complete, yet slow recovery. Another chemiresistive NO₂ sensor was developed by Prezioso *et al.*²¹⁹ by using pristine GO; this device yielded the best performance when heated at 150 °C, with a response of 60 % to 5 ppm of NO₂, yet accompanied by the response and recovery time close to an hour. In another work²²⁶ the authors improved this GO based NO₂ sensor which reached a detection limit of just 20 ppb. Furthermore, this sensor had a response to NO₂ unaffected by the humidity level and was able

to detect also the presence of ethanol, acetone and ammonia. Sinitskii *et al.*²²² fabricated a highly selective gas sensor able to discern three different alcohols such as methanol, ethanol and isopropanol with a high success rate by using a 20-channels array integrated sensor of thermally reduced GO, in which each channel had a unique response due to the structural/morphological irregularity of the rGO film.

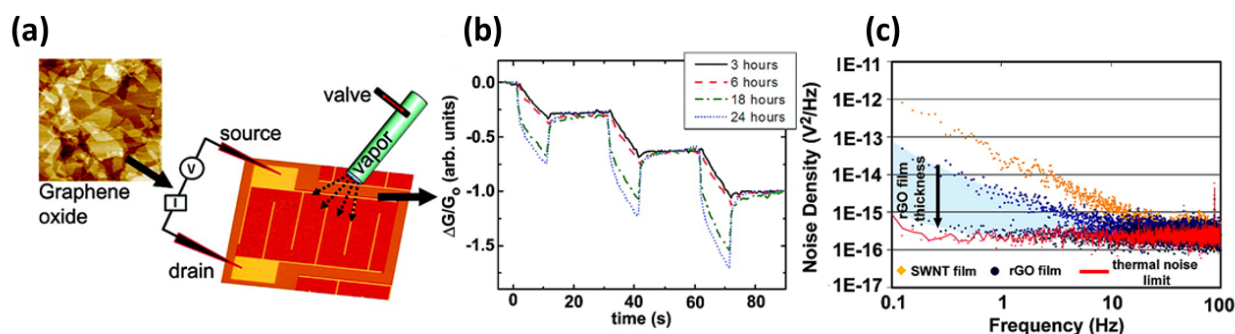


Figure 4. (a) AFM image of the GO film and scheme of the device. (b) Plot of the response to 5 s pulses of acetone for GO devices reduced with hydrazine for 3, 6, 18, and 24 h. (c) Noise density spectrum for SWNT and rGO devices. Adapted from Ref.⁷³ with permission from American Chemical Society.

Among the various gas sensing applications, GO and rGO have been extensively used also as humidity sensors.²²⁷⁻²³⁰ Humidity is the presence of water vapour in air and is normally measured as relative humidity (RH) which is the ratio between the partial pressure of water and the equilibrium vapour pressure at a given temperature. It is worth reminding that having a good control over the humidity in the environment is key for numerous industrial and technological processes and most importantly to ensure the comfort of live beings, improving the quality of life in living and working places

The presence of numerous oxygen-bearing functional groups as hydroxyl, epoxy and carboxylic groups renders GO's surface highly hydrophilic, thus being a perfect candidate for sensing moisture.⁷¹ Ruoff *et al.*²³¹ in 2008 demonstrated that the interaction of water molecules with these functional groups of GO determines changes in the conductivity of GO; the authors therefore concluded that highly reduced GO is not suitable for humidity sensing. Water interacts with GO mainly by increasing its conductivity because of the formation of charge carriers as hydronium and hydroxide ions as demonstrated by Yao *et al.* in 2012.²³⁰ In their work, GO was drop-cast on an interdigitated electrode (IDE); when a potentials $< 2V$

was applied a 10-fold change in the GO's conductivity was observed upon varying the RH between 15 and 95 %. Such a conductivity change was found being even larger when higher potentials were applied and has been attributed to the favoured ionization of the water molecules. Similarly Borini *et al.*²²⁷ spray-cast a solution of GO on top of interdigitated electrodes (IDE) screen-printed on a flexible and transparent polyacrylonitrile (PAN) substrate. Thin films of spray-coated GO can be essentially reduced to a simple resistor–capacitor (RC) equivalent circuit in which a resistor and a capacitor are in parallel. It was found that resistance displayed an impressive exponential dependence from RH, with a 10-fold decrease when increasing the RH from 40 to 70 % (Figure 5a). Furthermore, this device showed an ultra-fast response and recovery to pulses of humid air of just 20 ms (Figure 5b). Another GO based flexible humidity sensor was prepared by Guo *et al.*²²⁰ via spin-coating on PET substrates. The GO film was micro-patterned by using two-beam-laser interference to create alternating lines of GO and rGO. By tuning the laser power, it was possible to control the conductivity as well as the response/recovery time. The best results were obtained at 0.2 W laser power, with a response and recovery time of a few seconds and a change in the resistance of 1 order of magnitude between 11 and 95 % RH. In a very recent work, Tai *et al.*²²⁹ underlined the effect of the degree of reduction on the humidity sensing properties of GO. While GO showed a decrease in the resistance for increasing humidity, rGO displayed the opposite behaviour. By annealing GO at 150 °C for different times, integrating the so-obtained material in humidity sensors, and studying them with electrochemical impedance spectroscopy (EIS) they concluded that the resulting resistance of macroscopic film of GO is essentially composed by two resistances in series: an intrinsic resistance within the flake (R_{int}) and a junction resistance between flakes (R_{junct}). Both these terms are humidity dependent; in particular, the first can be expressed as two resistors in parallel, one in-plate resistance dependent only on the reduction grade and the ionic resistance, which decreases with the increasing of humidity. The second term R_{junct} can also be expressed as two elements in parallel, a capacitor (C_{EDL}) and a resistor ($R_{Ohm-contact}$) both humidity dependent: when humidity increases and the GO film adsorbs water molecules the distance between flakes increases, thus yielding to an increase in both the values C_{EDL} and $R_{Ohm-contact}$. As a result, with the increasing of humidity, R_{int} decreases (negative response) and R_{junct} increases (positive response). For GO R_{int} is greater than R_{junct} , thus leading to a negative response to the increased humidity, while for rGO $R_{int} < R_{junct}$ thus yielding to a positive response to the increased humidity (Figure 5d).

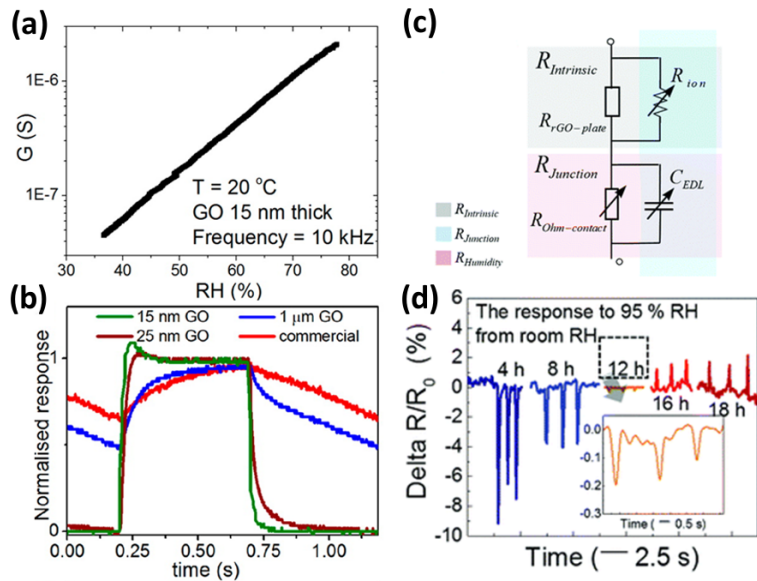


Figure 5. (a) Conductance dependence of an ultrathin GO film to the RH of the environment. (b) Normalized response of the different sensors to a modulated humid air flow at 1Hz. Adapted from Ref. ²²⁷ with permission of American Chemical Society. (c) Equivalent electrical circuitry model of rGO representing intrinsic and junction-dependent resistance. (d) Response of rGO films with different thermal reduction times to humidity pulses. Adapted from Ref. ²²⁹ with permission from the Royal Society of Chemistry.

In order to target more specific types of gases, GO and rGO have been functionalized with (macro)molecules bearing specific functional groups acting as selective active sites. The molecules can be chemisorbed on the surface of graphene or physisorbed on its surface. For example Al-Mashat *et al.* ²³² reported an hydrogen sensor based on an assembly of polyaniline (PANI) adsorbed on the rGO surface. The material was synthesized via ultrasonication of an alcoholic mixture of rGO, aniline and the polymerization initiators, which promoted the polymerization of PANI selectively on the rGO surface. Then, the mixture was spray coated on a quartz substrate with gold IDE. This functionalized rGO showed a better sensitivity to H₂ compared to pure rGO and PANI with a 16.6 % response to 1 % H₂ gas. The improved sensitivity was attributed to the higher porosity of the hybrid structure in respect to the bare PANI. In another case, an assembly PANI-rGO was used for NH₃ sensing ²³³. Shi *et al.* ²³⁴ fabricated a NO₂ chemresistive-type gas sensor based on sulfonated rGO (S-rGO) and rGO functionalized with ethylenediamine (EDA-rGO). In particular, S-rGO displayed a response to NO₂ 10 times greater of rGO and 3 times larger than EDA-rGO. Furthermore, the S-rGO exhibited also faster response and recovery and good selectivity towards NO₂. The superior sensitivity of the S-rGO sensors can be attributed

to the electron withdrawing characteristics of sulfophenyl groups, which enhance the hole doping in *p*-type rGO. After the absorption of NO₂ the *p*-doping is further boosted therefore harnessing the material's conductivity. Hu *et al.*²²³ developed a DMMP sensor based on rGO functionalized with *p*-phenylenediamine, which was deposited from solution on a IDE. The response of this device to DMMP vapours was much higher of the one of hydrazine reduced GO ($\Delta R/R_0$ to 30 ppm DMMP = 11 % for f-rGO and 1.5 % for rGO); however, the response and recovery times were on the scale of several minutes. The chemical functionalization of GO can be used to build sensor arrays in order to recognize different gases. In a very recent work Jelinek *et al.*²⁰⁰ assembled a capacitive-type porous graphene oxide (pGO) vapour sensor in which the GO was deposited on a IDE through a freeze-drying technique in order to maximize the area exposed to the gas. In parallel the GO was also chemically functionalized with aniline (phenyl-GO), dodecylamine (dodecyl-GO) and ethanolamine (ethanol-GO) by exploiting the amidation reaction between the amines and the carboxylic groups present on the GO surface – yet, the functionalization of epoxy groups may also occur. These three differently functionalized GOs were deposited on IDE similarly to pGO and by combining them in an array it was possible to selectively detect the presence of water, NH₃, toluene, EtOH, phenol and cyclohexane (Figure 6). Additionally, the sensitivities of these devices towards the target gases were high, with high reproducibility and both response and recovery times on the time scale of few seconds.

The doping of GO and rGO has also been used to improve the performance of humidity sensors. Rathi and Pal²²⁸ demonstrated that GO doped with Li and B have higher sensitivity to the humidity compared to GO. The Li-GO and B-GO were prepared through reaction of GO with LiOH and boric acid, then these doped GO water solutions were drop-casted on a glass with two copper electrodes on top. The Li doped GO in particular showed a response to the humidity 4 times higher than un-doped GO. Also in this case the higher performances can be attributed to the increased *p*-doping of Li-GO.

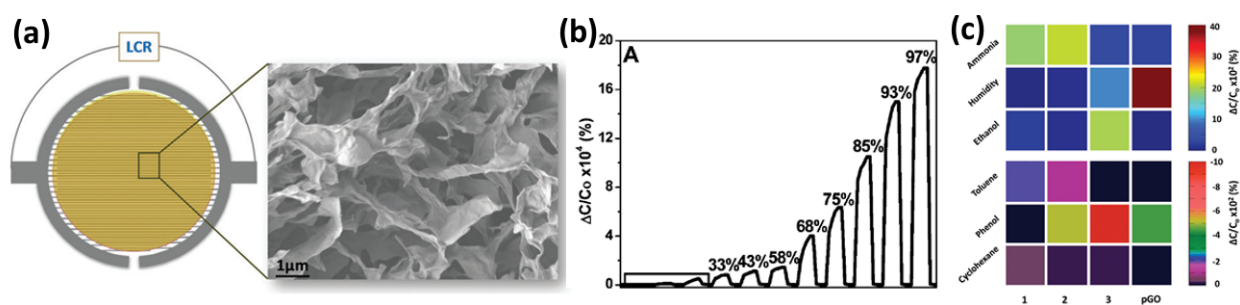


Figure 6. (a) Schematic structure of the device and SEM image of the p-GO. (b) Response of the p-GO device to different values of humidity (RH %). (c) Array-based colour code identification of different vapours (180 ppm) using the functionalized pGO capacitive sensor: phenyl-GO (1); dodecyl-GO (2); ethanol-GO (3). The colours indicated in the diagram correspond to the percentage capacitance response, according to the colour keys shown on the right. Adapted from Ref. ²⁰⁰ with permission from Royal Society of Chemistry.

2.1.3 Gas sensing with graphene based composites

Hybrid composites, made by combining graphene with functional materials (such as metals, metal oxides, polymers, *etc.*) can exhibit significantly increased porosity compared to films of pristine 2DMs, thus resulting in increased accessible surface area. Furthermore, such approach harnesses the sensing performances of 2DMs due to the combination of different functionalities and their synergistic effects.⁷⁵ Metallic nanoparticles (NPs) and films (especially those of noble metals) can enhance the sensitivity and selectivity due to specific catalytic effects.²³⁵

For example, it is well-known that the metals such as palladium and platinum exhibit a catalytic activity towards the adsorption of hydrogen.²³⁶ With this in mind, palladium and platinum-graphene composites have been extensively explored to target H₂ gas;^{182, 235, 237-241} hydrogen is colourless and odourless but mixed with air forms an explosive mixture; thus, its detection is of paramount importance for safety. Kaniyoor *et al.*²³⁵ produced a hydrogen sensor based on GO decorated with Pt NPs: the Pt NPs were grown *in situ* on the GO surface and then the so decorated GO was drop-casted on an alumina substrate with pre-patterned Cu electrodes. This device showed a 16 % increased electrical resistance when exposed to 4 % H₂ yet accompanied with a slow responsivity. Similarly, Wu *et al.*²³⁷ fabricated an H₂ sensor based on a Pd thin film (1 nm) evaporated over CVD graphene with 12 % sensitivity towards

the exposure to 1 % H₂. Chung *et al.*²³⁸ realized a flexible hydrogen sensor composed by CVD graphene decorated with Pd NPs exhibiting an impressive sensitivity of 30 % to 1 % H₂. Reduced GO decorated with Pd NPs has also been used to sense NO gas for medical applications by Li *et al.*¹⁸⁵ The rGO with physically absorbed Pd NPs on its surface was deposited on CVD graphene-coated Ni electrodes (

Figure 7a). This device exhibited high sensitivity with an impressive limit of detection of 2 ppb, being notably higher than the reference device assembled without decorated rGO or without graphene-coated electrodes (

Figure 7b). In this case Pd NPs may act as an absorption site of NO molecules, promoting the donation of electrons from NO to rGO. Furthermore, the interaction of the hybrid with NO can lower the Schottky barriers between rGO and Pd NPs leading to a further increase in the conductance.

Interestingly, the modification of rGO with functional groups capable of boosting the *p*-doping can be combined with the catalytic activity of metal NPs. For example Huang *et al.*²⁴² assembled a flexible, all-printed NO₂ sensor based on sulfonated rGO/Ag NPs. The rGO was firstly sulfonated and then Ag NPs were grown *in situ* on the rGO surface *via* hydrazine reduction of AgNO₃ (

Figure 7c). Then Ag electrodes and the rGO/Ag NPs based ink were printed on a PI substrate. This sensor displayed a 74.6 % fast response to 50 ppm of NO₂. Besides, the sensor was quite robust: it exhibited a good stability over time and tolerance to humidity (

Figure 7d). The sensor could also detect NH₃, with a similar, yet a negative response.

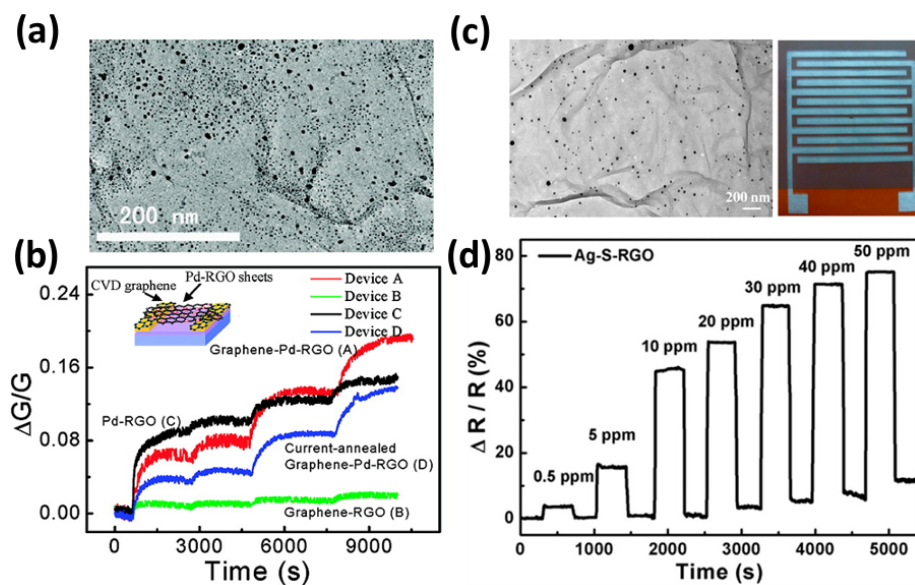


Figure 7. (a) Transmission Electron Microscopy (TEM) images of Pd-rGO sheets. (b) Relative changes in conductance of the various devices versus time when exposed to 1, 10, 60, and 120 s pulses of NO gas. Inset: Scheme of the device. Adapted from Ref. ¹⁸⁵ with permission from American Chemical Society. (c) TEM image of Ag-S-rGO and photographs of the sensing device printed on PI. (d) Response of the Ag-S-rGO sensor as a function of time in various concentrations of NO₂ gas. Adapted from Ref. ²⁴² with permission from American Chemical Society.

Hybrid materials composed by graphene and metal oxides nanostructures ^{184, 243-249} or quantum dots ^{250, 251} have also proved to enhance the sensing performances of graphene to NO₂, ethanol, H₂S and other gases. In fact, semiconducting metal oxides such as ZnO, SnO₂, In₂O₃, and Cu₂O, have been already employed in gas sensing, ^{252, 253} yet the high operating temperature and low conductivity represented a severe limit towards their application. Deng *et al.* ²⁴³ synthesized a hybrid rGO/Cu₂O nanowire mesocrystals under hydrothermal conditions and integrated it in a NO₂ gas sensor. The hybrid sensor displayed a higher sensitivity compared to that of the single components alone, with an impressive detection limit of 64 ppb (Figure 8a-b). Also in this case the response resulted from the electron withdrawing effect of NO₂ towards the *p*-type rGO and Cu₂O semiconducting mesocrystals, and it took further advantage of the higher porosity of the hybrid material. An ethanol sensor was assembled by Yi *et al.* ²⁴⁴ growing semiconductive vertically aligned ZnO nanorods on a metal bottom electrode and depositing CVD few-layers graphene on top as top contact. This sensor, with

geometry similar to metal-insulator-metal devices, offered a 900 % sensitivity to 10 ppm of ethanol and a good flexibility. However, an operating temperature as high as 300 °C was required to leverage the *n*-type characteristic of semiconductive ZnO nanorods and obtain a good sensitivity to the electron donor EtOH. In another work²⁴⁵ GO decorated with ZnO NPs was used to sense CO and NH₃ gases. ZnO NPs were grown in solution on the GO surface; such a decorated GO was spin-coated on ITO patterned glass substrates. This device, when operated at room temperature, revealed a much higher sensitivity towards CO and NH₃ compared with the two components alone. In particular, the change in conductivity was ca. 24 % to 1 ppm of NH₃ and 22 ppm of CO. Similarly, Cuong *et al.*²⁴⁶ reported a H₂S sensor based on vertically aligned ZnO nanorods hydrothermally grown on the surface of spray-coated rGO which showed high sensitivity in oxygen, but long response and recovery times. Another H₂S sensor was developed by Zhou *et al.*²⁴⁷ by hydrothermally growing Cu₂O nanocrystals onto the GO surface and drop-casting the suspension onto a gold IDE on Si/SiO₂ substrate. The GO/Cu₂O assembly was demonstrated to have better performances compared to the isolated components, with a high 11 % sensitivity to 5 ppb of H₂S at room temperature and responses and recovery times on the few minutes time scale. Assembly of rGO/In₂O₃ have been used to detect NO₂^{249, 254} for environmental monitoring. Feng *et al.*²⁵⁴ were able to embed In₂O₃ nanocubes in rGO networks using InN NWs and GO as precursors. Depending on the ratio between the two components, different morphologies were obtained: the best sensing performances were found when a 1:1 InN:GO ratio was used (Figure 8c). Such a sensor exhibited a 61 % change in the resistance when exposed to 5 ppm of NO₂ and a markedly high selectivity towards interfering gases (Figure 8d). In 2017 Liu *et al.*²⁴⁹ overtook these performances by combining flower-shaped In₂O₃ nanocrystals with rGO thereby obtaining a sensor with an impressive 3 orders of magnitude increase in the resistance when exposed to 1 ppm of NO₂ combined with a detection limit lower than 10 ppb.

The sensing of simple alkanes, as methane and ethane is important because of the industrial relevance of these molecules. However, very few examples of sensing of these gases can be found in literature because of the weak doping nature of these apolar molecules.²⁵⁵⁻²⁵⁷ For example a hydrogen and liquid petroleum gas (LPG) sensor based on rGO decorated with SnO₂ quantum dots has been developed by Lee *et al.*²⁵¹, showing an impressive response of 89 % to 500 ppm of H₂ and 92 % to 500 ppm of LPG. Zhang *et al.*²⁵⁵ produced a methane sensor based on a nanocomposite of rGO and ZnO nanocrystals that displayed a decent and fast response to 100 ppm of CH₄ and a good selectivity towards interfering gases;

unfortunately a high operating temperature of 190 °C is necessary to properly function this sensor.

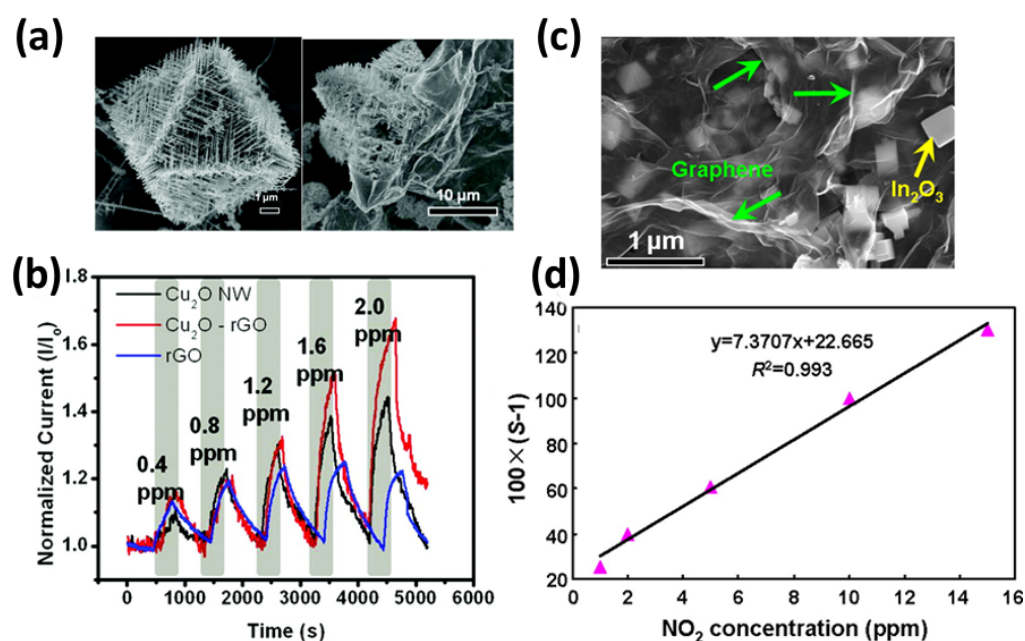


Figure 8. (a) SEM images of rGO-Cu₂O mesocystal composites, and (b) dynamic response of Cu₂O NW, rGO-Cu₂O, and rGO devices under increasing NO₂ exposure. Adapted from Ref. ²⁴³ with permission from American Chemical Society. (c) SEM images of the In₂O₃ cubes/rGO composites at 1:1 mass ratio. (d) Linear fitting curve of the sensor response versus NO₂ concentration. Adapted from Ref. ²⁵⁴ with permission from American Chemical Society.

Polymers and semiconducting polymers coupled with graphene, GO and rGO constitute the third large family of hybrid assemblies. Some of these assemblies with conducting polymers, like polyaniline/GO assembly, have already been discussed in the previous paragraph²³² as the boundary between polymer functionalized GO/rGO and polymer-GO/rGO assemblies is blurred. In a typical example, in 2014 Zhang *et al.*²²¹ developed a resistive-type, flexible humidity sensor based on the assembly of GO and a polyelectrolyte, i.e. poly(diallylimethylammonium chloride) (PDDA), by using layer-by-layer (LbL) deposition method. The nanostructured film was fabricated on a Cu/Ni IDE on PI substrate. In first instance a bi-layer of PDDA and an ionic polymer was self-assembled on the substrate as precursor layers for charge enhancement, and then 5 GO/PDDA layers were deposited by alternating immersions into PDDA and GO suspensions for five repetitive cycles. Following, the GO was chemically reduced by soaking the film in solution of HBr acid. Such an assembled device exhibited a 37 % increase in the resistance when passing from 0 % RH to

97 % RH, with a response and recovery time of a few minutes, an excellent linearity of the response and a long stability over time. The superior sensor performances were attributed to high surface area of the assembly, the *p*-type semiconducting properties of rGO at low RH, and to the interlayer swelling of PDDA/RGO film at high RH, which contributes to the resistance increase. In another work Shi *et al.*²⁵⁸ synthesized GO/conducting polymer composite hydrogels, including GO/polypyrrole (PPy) and GO/PANI, by chemical polymerization *in situ* of the corresponding monomers in aqueous dispersions of GO, where GO sheets acted as 2D template. These gels showed many interesting properties, including electrical conductivity and electrochemical activity. In particular, the GO/PPy gel, once lyophilized, exhibited a good sensitivity to NH₃ gas. In this case, the improved response was attributed to the high surface area of the conductive gel compared to the bare polymer film.

A nanocomposite thin film of chemically exfoliated graphene and PANI was used by Wu *et al.*²⁵⁷ to sense methane at room temperature, with a decent 10 % response to 10 ppm of gas. Shi *et al.*²⁵⁹ prepared a highly sensitive NO₂ sensor based on electrospun PVA/PEI nanofibers coated with a thin GO layer, which self-assembled on the surface of the nanofibers for charge interaction. The nanofibers were deposited on an IDE electrode and then exposed to hydrazine vapours to reduce the GO. This sensor displayed an impressive sensitivity, with a 16 % response to 150 ppb of NO₂ and good linearity until 1 ppm; furthermore, the stability and the selectivity towards NO₂ were high, and the response and recovery time were around 5 minutes.

The combination of graphene with ionic liquid is a similar route to achieve high sensitivity and selectivity in gas sensing. Within this framework, Ariga *et al.*²⁶⁰ developed a sensor capable of discriminating between various organic vapours based on a multi-layered rGO/ionic liquid film assembled in solution *via* LbL method on a quartz microbalance. Assemblies of CNTs and graphene have been also employed for gas sensing. In 2010 Jeong *et al.*²⁶¹ produced a flexible NO₂ sensor using a CNTs/rGO hybrid film. A rGO film was deposited on a plastic substrate with a gold IDE followed by the CVD deposition of vertically aligned CNTs on top. The so-fabricated sensor revealed a sensitivity of 20% after 60 min exposure to 10 ppm NO₂ at room temperature.

The most importing performance parameters of all the graphene based gas sensors described in this section, including the response time and the sensitivity (calculated as the ratio between

the response % and the corresponding gas concentration expressed in ppm) are summarized in Table 1.

Table 1. Comparison of the sensing performances of different graphene based gas sensors.

<i>Material</i>	<i>Sensed gases</i>	<i>Response/recovery time</i>	<i>Limit of detection</i>	<i>Sensitivity (Response * ppm⁻¹)</i>	<i>Ref.</i>
<i>SLG (ME)</i>	NO ₂	500 s	ppb	4%	201
	H ₂ O, nonanal, octanoic acid		0.5 ppm	0.6 - 1 %	204
	CO ₂	10 min/10 min		0.26 %	206
	NO ₂	4 min/4 min	100 ppm	0.08 %	80
<i>SLG/BLG (ME)</i>	THF, MeOH, MeCN, CHCl ₃				205
<i>SLG (CVD)</i>	CH ₄ , NH ₃ , H ₂	9 min/9 min		0.046 % (NH ₃)	211
<i>SLG/BLG (CVD)</i>	NO ₂	95 s/11s		0.98 %	209
<i>SLG(CVD) oxidized with O₃</i>	NO ₂	15 min/30 min		0.0085 %	208
<i>SLG (CVD) patterned</i>	NO ₂ , NH ₃	89 s/579s		2.6 % (NO ₂)	210
<i>CVD graphene Foam</i>	NH ₃	800 s/1200 s		0.03 %	212
<i>SLG (epitaxially grown)</i>	NO ₂	1 h/3 h		48%	216
	NO ₂	250 s/150 s at 300 °C		0.55 %	214
<i>GO</i>	H ₂ O	20 ms/20 ms		30 %/RH	227
	NO ₂ , H ₂ O	40 min/40 min		12%	219
	NO ₂	40 min/40 min	20 ppb	250 %	226
	H ₂ O			12.5 %	230
<i>rGO</i>	Acetone, DMMP, HCN	5 s		100%	73
	NO ₂ , Cl ₂ , NH ₃	10 min/10 min		0.25 % (NH ₃)	224
	H ₂ O				231
	NO ₂	10 min/30 min		0.05 %	225
	H ₂ O	20 s/30 s		10.7 %/RH	148
<i>Li and B doped GO</i>	H ₂ O	4s		35.3 %	228
<i>Sulfonated-GO and EDA-GO</i>	NO ₂			50%	234
<i>phenyl-GO, dodecyl-GO ; ethanol-GO</i>	Various gases	15 s/10 s		200% (NH ₃)	200
<i>p-phenylenediamine/GO</i>	DMMP	1080 s/360 s		0.4 %	223

<i>PANI/rGO</i>	H ₂	2 min/3 min		0.0016 %	232
<i>PANI/Graphene flakes</i>	CH ₄	1 min/1 min	10 ppm	2.5 %	257
<i>PDDA/GO LbL assembly</i>	H ₂ O	2 min/3 min		0.38 %/RH	221
<i>microfiber PVA-PEI/rGO</i>	NO ₂	4 min/10 min		110%	259
<i>CNTs/GO</i>	NO ₂	1 h/3 h	0.5 ppm	2%	261
<i>Ag NPs/Sulfonated-GO</i>	NO ₂ , NH ₃	12 s/20 s		1.49 % (NO ₂)	242
<i>Pd film/CVD SLG</i>	H ₂	>10 min		0.0012 %	237
<i>Pd NPs/CVD SLG</i>	H ₂	15 min/20 min		0.003 %	238
<i>Pd/rGO</i>	NO	265 s		35%	185
<i>Pt NPs/GO</i>	H ₂	9 min/20 min		0.004 %	235
<i>Cu₂O NPs/rGO</i>	H ₂ S	5 min/5 min		2200%	247
<i>Cu₂O/rGO</i>	N ₂ O	5 min/10 min		34%	243
<i>In₂O₃/rGO</i>	NO ₂	4 min/1 min		109800%	249
<i>In₂O₃/rGO</i>	NO ₂	3 min/4 min		12.2 %	254
<i>ZnO NPs/rGO</i>	CO, NH ₃	5 min		24 % (NH ₃)	245
<i>ZnO Nrods/rGO</i>	H ₂ S	>30 min			246
<i>ZnO Nrods/CVD SLG</i>	EtOH			90%	244
<i>ZnO NCryst/rGO</i>	CH ₄	1 min/10 s	100 ppm	0.05 %	255

List of abbreviations. SLG : single layer graphene, ME : mechanically exfoliated, BLG : Bi-layer graphene, CVD : chemical vapour deposition, GO : graphene oxide, rGO : reduced graphene oxide, CHCl₃ : chloroform, THF : tetrahydrofuran, MeOH : methanol, MeCN: acetonitrile, DMMP : dimethyl methylphosphonate, HCN : hydrogen cyanide, EDA : ethylenediamine, PANI : polyaniline, PDDA : poly(diallylimethylammonium chloride), LbL : layer-by-layer, PVA : polyvinyl alcohol, PEI : polyethylenimine, CNTs : carbon nanotubes, NPs : nanoparticles, Nrods : nanorods, NCryst: nanocrystals.

2.2 Gas sensing with transition metal dichalcogenides

The sensing properties of TMDs are based on the already discussed charge-transfer mechanism. Yue *et al.*¹⁹¹ attained a deep insight into such a mechanism by focusing their attention to the *n*-type MoS₂ as exemplary system. In this article, the charge transfer mechanism between different gas molecules including O₂, H₂O, NH₃, NO, NO₂, CO, etc. and monolayer MoS₂ was explained with the aid of DFT calculations. The authors showed that the conduction band (CB) of pristine *n*-type MoS₂ monolayers is already populated by some electrons at room temperature. When the monolayer is exposed to electron-acceptor gases such as O₂, H₂O, NO, NO₂, and CO, electrons are transferred from MoS₂ to the sensitive gases, leading to a decrease of carrier density in MoS₂ which ultimately yields an increased

resistance. On the contrary NH_3 , which behaves as an electron-donor, transfers electrons to the MoS_2 monolayer thereby reducing its resistance.

The charge transfer between TMDs monolayers and gases can also be evidenced by changes in the photoluminescence (PL), which is due to the direct band gap properties of TMDs monolayers. Tongay *et al.*²⁶² showed that the light emission efficiency of these TMDs can be modulated by physisorption of gas molecules such as O_2 and H_2O , as a result of a molecular gating effect. The charge depletion in *n*-type materials such as MoS_2 and MoSe_2 caused by the charge transfer to molecules such as O_2 and H_2O leads to the drastic enhancement in photoluminescence for the stabilization of neutral excitons X_0 .

Among the various TMDs, MoS_2 is the most studied material. Mechanically exfoliated MoS_2 has been successfully used to detect a wide range of gases.^{14, 190, 193, 263} For instance Zhang *et al.*¹⁹⁰ fabricated miniaturized FETs based on mechanically exfoliated MoS_2 for the detection of NO. Single, bi, tri and quadri-layers of MoS_2 transferred on Si/SiO₂ substrates were investigated in this work. The sensitivity of these devices to NO was high and the response fast; however, the current passing through a single-layer thick MoS_2 FET was unstable while the devices based on 2, 3 and 4-layers MoS_2 showed both high sensitivity and stability with a detection limit of 0.8 ppm of NO. In a similar work Late *et al.*¹⁴ prepared transistor depositing mechanically exfoliated MoS_2 flakes on Si/SiO₂ substrates and contacting them with gold top electrodes through electron beam lithography and they employed them to sense NO_2 , NH_3 and humidity (Figure 9a). From the analysis of flakes consisting of 1, 2 and 5-layers it was observed that the signal was unstable for the case of monolayers while among 2 and 5-layers the latter exhibited the highest response to the sensed gases. As expected from the charge transfer mechanism NO_2 produced a negative resistance response while NH_3 had the opposite behaviour (Figure 9b and 8c). For the same reason the applied gate voltage had opposite effects in NH_3 and NO_2 in increasing or decreasing the sensitivity and these effects were larger for the 5-layers device. In addition, the effects of green light irradiation over the sensitivity, and response time of the device was also explored. The reason for the increased and stable response in the devices based on few-layers compared to the single-layer thick remains unclear, yet the authors showed with DFT calculations that the adsorption of NO_2 was slightly favoured in the few-layers thick devices. In another example Perkins *et al.*¹⁹³ measured the conductance changes in a device based on mechanically exfoliated MoS_2 after exposing it to various organic vapours. They found out that the device had high selectivity

towards triethylamine (TEA) and acetone with a detection limit of 1 ppm over the last analyte.

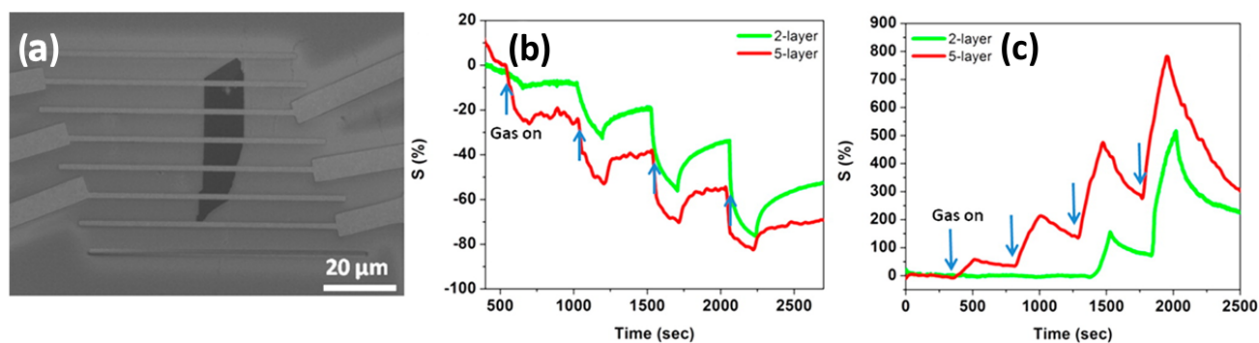


Figure 9. (a) SEM image of two-layer MoS₂ transistor device. Sensing performances of the 2 and 5-layers devices when exposed to 100, 200, 500, 1000 ppm of (b) NH₃, and (c) NO₂. Adapted from Ref. ¹⁴ with permission from the American Chemical Society.

Dispersion of MoS₂ sheets in solution can be obtained with a relatively high yield and low costs by liquid-phase exfoliation. These dispersions can be processed to generate films of overlapping flakes, which have been used as active materials to detect different gases.^{192, 264, 265} For example Yao *et al.*²⁶⁵ prepared a water dispersion of MoS₂ flakes combining grinding and sonication of the bulk material. The dispersion was deposited with an inject printer onto pre-patterned microelectrodes and the so-obtained device was used to detect NH₃ gas up to a few ppm. A similar method that relies on the intercalation of Li ions within the MoS₂ layers and the subsequent exfoliation of these in water solution through ultrasonication has been used by Zhang *et al.*¹⁴¹ to sense NO. The dispersion of MoS₂ flakes was deposited by drop-casting on a Si/SiO₂ substrate pre-patterned with gold electrodes to form a thin-film transistor (TFT) which was found being sensitive to NO in concentrations down to 0.4 ppm.

Similarly to graphene, also MoS₂ can be produced by CVD.¹⁴³ In 2014 Liu *et al.*¹⁴⁴ fabricated FETs by contacting through EBL small triangular-shaped monolayers of MoS₂ grown by CVD on Si/SiO₂ substrates (

Figure 10a). These devices exhibited a high Schottky barrier (ΔSB) between the Ti/Au contacts and the monolayer, which can be modulated by exposing the devices to NO₂ and NH₃ gases. In particular, NO₂ adsorption increases the ΔSB and thus of the contact resistance, whereas the NH₃ adsorption determines a decrease of both them. The effect of the Schottky

barrier modulation, together with the charge transfer mechanism, yielded an increase in the sensitivity of the device to both the analytes, with changes in the conductance up to 3 orders of magnitude and detection limits down to 20 ppb of NO₂ and 1 ppm of NH₃ (

Figure 10b). Cho *et al.*¹⁴⁶ produced a gas and photo-detector based on wafer-scale film of 3-layers MoS₂ deposited on sapphire substrates *via* CVD with IDE gold electrodes on top. This sensor exhibited good sensitivity and selectivity to NO₂, with a response around 40 % to 120 ppb combined with a stable cycling behaviour. Furthermore, the response time was fast, yet the recovery at room temperature was not complete even after 30 minutes. Similarly, a 3-layers MoS₂ film produced by CVD on sapphire substrates was developed by Cho *et al.*²⁶⁶ for the sensing of NO₂ and NH₃. The device exhibited an increased resistance when exposed to NO₂ and a decreased resistance when interacting with NH₃, as expected from the charge-transfer mechanism.

The sensitivity of MoS₂ to gases could be improved by increasing the number of active sites that can interact with the analyte of interest through processes of molecular absorption on sulphur vacancies and edges. For example Jung *et al.*¹⁴⁷ showed that sensors based on MoS₂ layers vertically aligned, thus exposing their edges to the environment, exhibited a higher response to NO₂ and EtOH gases. The films of horizontally and vertically aligned MoS₂ layers were produced *via* CVD. By increasing the thickness of the films, the number of vertically aligned layers increased, and at a thickness of 15 nm almost all the surface was covered by vertically aligned layers, as revealed by TEM mapping. These films exhibited a response 5 times greater than the 1 nm thick film, because of the increased edge sites exposed to the environment.

The surface functionalization of MoS₂ with molecules or nanomaterials as metal and metal oxides NPs represents an alternative viable method to increase the interactions with the gas molecules, thus to harness the device sensitivity.²⁶⁷⁻²⁶⁸ For example Zhang *et al.*²⁶⁷ fabricated a flexible TFT depositing a film of MoS₂ obtained with the method previously described¹⁴¹ on an rGO electrode array on polyethylene terephthalate (PET). Successively, Pt NPs were deposited on the surface of the MoS₂ film and the device was employed to detect NO₂. The device functionalized with the Pt NPs showed a response to NO₂ being 3 times faster than the not-functionalized device, with detection limits as low as 2 ppb. The reason for such an improved sensitivity was attributed to the formation of a Schottky barrier between the MoS₂ and the Pt NPs. In a similar work Sarkar *et al.*²⁶⁹ functionalized the MoS₂

with various metallic NPs, included Ag, Pd, Pt, Sc and Y and examined their doping effect. The Pt NPs functionalized MoS₂ device revealed an improved sensitivity to H₂ gas compared to the pristine MoS₂ device. An H₂ sensor based on Pd NPs functionalized MoS₂ was assembled also by Jin *et al.*²⁶⁸ by drop-casting of a water suspension of MoS₂ flakes obtained by liquid-phase exfoliation and Pd NPs. This sensor displayed a fast response and good recovery upon exposure to of H₂ (0.5-5%) and a decent cross-sensitivity. Metal oxides NPs have also been used to improve the sensitivity; in particular, the semiconducting SnO₂ has been employed to sense humidity,¹⁹⁸ NO₂²⁷⁰ and EtOH.²⁷¹ For example Chen *et al.*²⁷⁰ prepared a NO₂ sensor based on SnO₂ nanocrystals decorated MoS₂ nanosheets. The material was prepared as depicted in

Figure 10c. A MoS₂ flakes suspension was produced by lithiation of MoS₂ followed by liquid exfoliation, then SnCl₄ was added at the suspension and Sn⁴⁺ ions got absorbed on the negatively charged flake's surface. Finally, the SnO₂/MoS₂ hybrid was obtained by annealing of the filtered suspension at 300 °C in argon. Interestingly, this material showed *p*-type behaviour because of the *p*-doping effect of the SnO₂ nanocrystals. The sensor based on the hybrid showed a decent sensitivity to NO₂, with a 30 % response to 10 ppm and detection limit of 0.5 ppm combined with a good reversibility and selectivity (

Figure 10d).

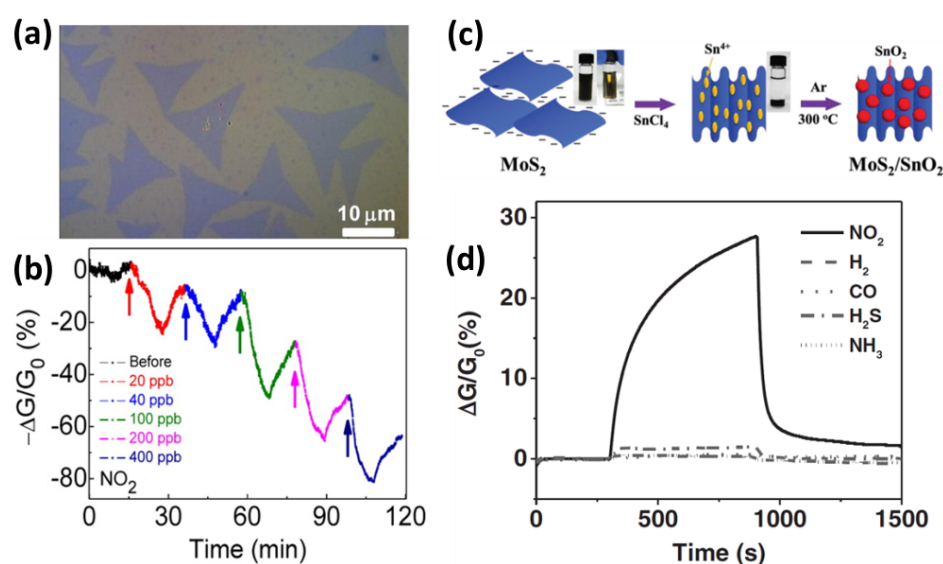


Figure 10. (a) Micrograph of the triangular shaped of CVD grown MoS₂ monolayers supported on a Si/SiO₂ substrate. (b) Response of the device under the exposure to increasing NO₂ concentrations. Adapted from Ref. ¹⁴⁴ with permission from American Chemical

Society. (d) Schematic illustration of the preparation of the SnO₂/MoS₂ hybrid. (d) Response of the device to 10 ppm of various gases, highlighting the selectivity towards NO₂, the sensitivity and the response time. Adapted from Ref. ²⁷⁰ with permission from Wiley-VHC.

MoS₂ can also be functionalized with dangling molecules, especially through the covalent attachment of thiolated molecules on the sulphur vacancies of the MoS₂ surface.²⁷² By applying such a strategy Jung *et al.*¹³ developed a volatile organic compounds (VOCs) sensor for breath analysis based on MoS₂ nanosheets obtained *via* liquid-exfoliation functionalized with mercaptoundecanoic acid (MUA). This sensor exhibited a good sensitivity down to 1 ppm to various VOCs included toluene, hexane, acetone, propanale and ethanol. Furthermore, while the resistance of pristine MoS₂ based device increased upon exposure to all the sensed gases, the MUA-MoS₂ sensor once exposed to the molecules with oxygen functionalities displayed a decrease in the resistance. These differences originated from the diverse interaction of the oxygen containing molecules with the dangling MUA. In particular, the oxygen-bearing gas molecules can interact *via* hydrogen bonding with the exposed carboxyl group in MUA and then form an electron-rich region able to transfer the charge to the MoS₂ sensing channel.

Among the other TMDs, MoSe₂, WS₂, WSe₂, ReS₂ and PtSe₂ have also been employed for fabricating gas sensing devices by exploiting their diverse semiconducting characteristics and their unique high surface-to-volume ratio. For example MoSe₂ has found application as NH₃ sensor in a work of Late *et al.*²⁷³ The device was based on a single-layer MoSe₂ flake obtained *via* mechanical exfoliation deposited on a Si/SiO₂ substrate and contacted with gold electrodes by EBL. The device showed a sensitivity to NH₃ in concentration from 50 to 500 ppm and a response time of 2.5 min. The doping effect of NH₃ was confirmed also by the shift in the Raman spectrum after the absorption of NH₃.

WS₂ has been used in recent years as a sensor for NH₃,^{52, 194} humidity^{195, 274} and alcohols.^{52, 195} Li *et al.*⁵² developed a transistor based on a mechanically exfoliated multilayer WS₂ flake and measured the current change and the photoresponse of the transistor when exposed to different gases included NH₃, ethanol and O₂. In particular, NH₃ and ethanol behaved as electron-donor, yielding an increase in current and photo-responsivity in the *n*-type WS₂ device, while O₂ extracted electrons from the device producing a current decrease and a lower photo-response. The device showed the highest sensitivity toward NH₃ and really fast response and recovery to all the sensed gases. Pumera *et al.*¹⁹⁵ prepared a chemiresistive

device depositing metallic 1T-WS₂ flakes obtained through liquid exfoliation on a IDE. The selective vapour sensing capabilities of the device were studied with impedance spectroscopy. When exposed to methanol the device presented an impedance phase spectrum with a resonant frequencies maximum around 1 Hz, while with water the maximum shifted to 1 kHz. Therefore, the device was able to discern methanol and water isolated and mixed together by simply selecting specific frequencies, even when high concentrations of interfering gases were present. Duesberg *et al.*¹⁹⁴ reported a NH₃ sensor based on CVD WS₂ thin film. The films with thickness ranging from 1 to 50 nm were grown on Si/SiO₂ substrates with gold IDE by using a H₂S plasma to sulphurise WO₃ films at 500 °C. The so-achieved devices demonstrated good sensitivity to NH₃, with a detection limit of 1.4 ppm.

WSe₂ is another TMD semiconductor which has been used to detect toxic gases. Javey *et al.*²⁷⁵ assembled a *p*-type FET based on mechanically exfoliated WSe₂ monolayer for NO₂ detection. A high work function metal as Pd was used as source and drain electrode to lower the contact resistance for holes injection. Upon exposure to 0.05 % of NO₂ the source-drain current increased of 5 magnitude orders as a result of the lowering of the Schottky barrier and increasing of *p*-doping. Furthermore, in devices with top-contact gate also the on-off ratio was greatly improved after NO₂ exposure.

Recently also ReS₂, ReSe₂ and PtSe₂ have been employed for gas sensing.²⁷⁶⁻²⁷⁹ Yan *et al.* reported the photoresponse of monolayer ReS₂²⁷⁶ and ReSe₂²⁷⁸ based FET under red light illumination in NH₃ and O₂ gases. The increased current was observed under illumination were higher in NH₃ atmosphere for ReS₂, while for ReSe₂ were higher in O₂ atmosphere. With the aid of DFT spin-polarized calculations the reason of the different photoresponse behaviour of single-layer ReS₂ and ReSe₂ FETs in O₂ and NH₃ was attributed to the different transfer direction and quantity of electrons in *n*-type ReS₂ and *p*-type ReSe₂. Yim *et al.*²⁷⁹ prepared a FET depositing a thin layer PtSe₂ film by CVD on a Si/SiO₂ substrate and tested it as a photodetector and gas sensor. The FET exhibited a fast detection of NO₂, with a resistance decrease of 0.25 % to 1 ppm of gas and response and recovery time of 30 and 10 s respectively.

2.3 Gas sensing with other 2DMs

SnS₂ is a layered material similar to TMDs exhibiting a higher electronegativity, which can potentially enhance the gas sensing capabilities. Ou *et al.*²⁸⁰ developed a sensitive and selective NO₂ sensor based on few-layer SnS₂ flakes obtained by a wet chemical synthesis technique. The flakes, with lateral size ranging between 80 and 200 nm, were synthesized from solution of SnCl₄ and sulphur and then drop-casted on an IDE. The device displayed a linear response from 0.6 to 10 ppm of NO₂, with an extremely high sensitivity of 3500 % to 10 ppm of NO₂ and response and recovery time of 170 and 140 s, respectively, at an operating temperature of 120 °C. Furthermore, the device showed high selectivity to NO₂ among high concentration of various interfering gases; these results were attributed to the strong affinity of SnS₂ for NO₂ as well as the favourable Fermi level of SnS₂ to the energy of the partially occupied NO₂ molecular orbitals.

In another example, Shi *et al.*²⁸¹ developed a NH₃ sensor operating at room temperature based on flower-shaped nanoflakes of SnS₂ prepared by a solution method. The sensing device was produced depositing a paste made of the interconnected flakes on a IDE. The sensor, when exposed to concentrations of NH₃ ranging between 5 and 150 ppm, presented a 21 % sensitivity to 5 ppm of gas with response and recovery time of 40 and 100 s, respectively. Another NH₃ sensor based on SnO₂/SnS₂ hybrid flakes was fabricated more recently by Xu *et al.*¹⁹⁷ To form the hybrid structure, SnS₂ flakes were firstly produced from a solution method and deposited on a gold IDE, then the device was heated to 300 °C in air in order to partially oxidize the surface to SnO₂. This device showed good sensitivity to NH₃, with a 16 % response to 10 ppm of NH₃, which was much higher compared to the sample with not oxidized SnS₂ or with the completely oxidized SnO₂. The device exhibited also a good dynamic range from 10 to 500 ppm of NH₃, high reproducibility and a response time of just 11 s.

Black phosphorus is a promising candidate as platform for gas sensing for its electrical properties and its sensitivity to the environments, which is usually considered as a drawback.¹⁵³ First-principle calculations¹⁵⁹ made it possible to cast light onto the interaction of many gases with phosphorene. The gas molecules can get physisorbed on the phosphorene surface and interact via van der Waals forces, yielding a modification of the electronic properties of phosphorene. O₂, NO₂ and NO act as strong electron acceptors whereas CO, H₂, H₂O and NH₃ act as electron donor. Abbas *et al.* firstly reported on the chemical sensing of NO₂ using FET based on mechanically exfoliated multilayer phosphorene.¹⁵⁸ The FET was produced by exfoliating the black phosphorus with the scotch-tape method on a Si/SiO₂ substrate and

patterning gold electrodes by EBL. The sensor displayed increased conductivity upon NO₂ exposure with a 3 % response to 5 ppb of the gas. Moreover, when the device was exposed to NO₂ concentrations ranging from 5 to 40 ppb, its response complied with the Langmuir isotherm for molecules adsorbed on a surface, confirming the charge transfer as sensing mechanism. Cui *et al.*⁹⁶ reported a similar FET based on phosphorene nanosheets exfoliated with the scotch-tape method for NO₂ sensing. The sensor showed thickness dependent performance with sensitivities up to 190 % to 20 ppb of NO₂, excellent stability and high selectivity in presence of interfering gases.

Phosphorene demonstrated good sensitivity also to organic vapors¹⁶⁰ and humidity.²⁸² Pumera *et al.*¹⁶⁰ developed a highly selective methanol sensor based on black phosphorus plates deposited on a gold IDE. The device was studied with EIS: when exposed to methanol a strong signal at a resonance frequency of 1 kHz appeared in the impedance phase spectrum and its intensity was proportional to the methanol concentration in the range of 380-1900 ppm. High reproducibility, long-term stability and excellent selectivity in presence of interfering gases were also established. Yasaei *et al.*²⁸² explored the possibility of employing of phosphorene films as humidity sensors. Phosphorene flakes were produced by ultrasonication in liquid and then filtered to obtain thick film of stacked flakes that were used as chemiresistors. When the humidity was increased from 10 to 85 % RH the conductivity increased by 4 orders of magnitude. According to EIS measurements it was hypothesized a sensing mechanism based on the formation of ionic charge carriers caused by ionization of the water molecules and solvation of the phosphorus oxoacids.

Hexagonal boron nitride possesses also promising properties, which render it an interesting material for gas sensing as it was assessed in few first-principles calculations.^{283, 284} However, the only example of application of hBN as gas sensor as been tendered by Feng *et al.*¹⁹⁹ In particular, a O₂ and CH₄ sensor based on wafer-scale hBN nanosheets (1.7 nm thick) by the CO₂-pulsed laser deposition technique was developed. Such sensor revealed sensitivity to 100 ppm O₂, of 150 % with response and recovery time around 70 s and 100 s, respectively. The sensitivity to 100 ppm CH₄, was even greater, being around 780 % with response and recovery times reduced to 15 s and 20 s, respectively. The highest sensitivity towards CH₄ could be caused by the higher polarizability compared to O₂ which promotes its adsorption on the surface and at grain boundaries of hBN, thereby changing its conductivity.

The properties and the performances of the gas sensing devices based on TMDs and other layered 2DMs are summarized in Table 2.

Table 2. Comparison of the sensing performances of different graphene based gas sensors.

<i>Material</i>	<i>Sensed gases</i>	<i>Response/ recovery time</i>	<i>Limit of detection</i>	<i>Sensitivity (Response* ppm⁻¹)</i>	<i>Ref.</i>
<i>MoS₂ (ME)</i>	NO		0.8 ppm		190
	NO ₂ , NH ₃ , H ₂ O			1.372% (NO ₂)	14
	TEA, acetone	15 s, 30 s	1 ppm TEA	5 % (TEA)	193
<i>MoS₂ (CVD)</i>	NO ₂ , NH ₃	> 10 min	5 ppm	2.5 % (NO ₂)	266
	NO ₂	1 min, 30 min	120 ppb	333%	146
<i>MoS₂ (CVD-Schottky contact)</i>	NO ₂ , NH ₃	10-20 min	20 ppb NO ₂ , 1 ppm NH ₃	1000% (NO ₂)	144
<i>MoS₂ (CVD - vertically align.)</i>	EtOH, NO ₂	10 min, 1h		0.1% (NO ₂)	147
<i>MoS₂ (LPE)</i>	NH ₃		ppm	5%	265
<i>MoS₂ (LPE in the presence of mercaptoundecanoic acid)</i>	VOCs (toluene, EtOH, hexane, acetone)			0.003-0.0015 %	13
<i>MoS₂- array of devices (LPE)</i>	NO ₂	1h, 1h	2 ppb	10%	267
<i>MoS₂/graphene electrodes</i>	NH ₃ , NO ₂		1.2 ppm NO ₂	0.06 % (NH ₃), 1.4 % (NO ₂)	263
<i>SnO₂/MoS₂</i>	NO ₂	5 min, 1 min	0.5 ppm	3%	270
	H ₂ O			34883%/RH	198
<i>Pd/MoS₂</i>	H ₂	40 s, 83 s		0.0032%	268
<i>WS₂</i>	H ₂ O, MeOH		5.6 ppm MeOH, 10 % RH		195
	NH ₃	ca. 10 min	1.4 ppm	0.005 %	194
<i>WSe₂</i>	NO ₂			200%	275
<i>PtSe₂</i>	NO ₂	30 s, 10 s		0.25 %	279
<i>SnS₂</i>	NO ₂	170 s, 140 s	30 ppb	350%	280
	NH ₃	11 s	10 ppm	1.6 %	197
<i>Phosphorene</i>	NH ₃	40 s, 100 s		4.2 %	281
	NO ₂	30 s, 840 s		600%	158
<i>hBN</i>	H ₂ O			135 %/RH	282
	O ₂ , CH ₄	15 s, 20 s		7.8 % (CH ₄)	199

List of abbreviations. ME : mechanically exfoliated, TEA : triethylamine, CVD : chemical vapour deposition, align. : aligned, EtOH : ethanol, LPE : liquid phase exfoliation, MeOH : methanol.

By and large, in this first part of this Review article we have provided an overview on the recent developments in the application of the most common 2DMs as gas sensors. From the results highlighted it is possible to affirm that 2DMs are fully competitive materials for production of high-performance gas sensors with low operating temperature thanks to their peculiar properties. Among 2DMs, pristine graphene does not appear being ideal for assembling sensing devices because of its chemical inertness and zero-band gap, and thus lack of semiconducting properties. In fact, the modulation of the semiconductivity is a key characteristic to harness the performances of gas sensors. Furthermore, in pristine graphene the absence of dangling groups on its surface that act as bonding sites for the analytes hampers the gas absorption and the sensing performances. Towards this end, doping and modifying graphene with functional atoms and groups and assembly hybrid structures of graphene and functional nanomaterials have been proven to be the winning strategy for enhancement of gas sensing performance. A different approach relies on exploiting other 2DMs, which possess intrinsic functionality as TMDs (MoS_2 , MoSe_2 , WS_2 etc.). Unlike graphene, TMDs show remarkable semiconducting properties with a tuneable band gap, which depends on their thickness and on the doping effect of *ad hoc* molecules as the sensed gases. Therefore, TMDs are overall promising materials for the fabrication of gas sensors. Nevertheless, up to date, their (production) costs and the hurdle in the control over the thickness hinder their application in competitive gas sensing devices.

A few challenges still need to be tackled in order to make 2DMs competitive for the industrial sensing market. In particular, while the sensitivity has evidenced great improvement, an enhancement of response and recovery characteristics is necessary. In fact, in most of the reported examples 2DMs based gas sensors exhibit slow response and incomplete recovery when operating at room temperature. Additional performance parameters that require improvements are the selectivity and stability. Most of these materials show a poor selectivity for a cross-response to different sensed gases and in many cases the response signals are unstable. Such a modest selectivity can hardly find a definitive solution because of the major hurdle of supramolecular chemistry to develop highly selective receptors for volatile gas molecules. Finally, the fabrication methods often used as microfabrication and drop casting are not suitable for the industrial scale up for their complexity and time-demand.

3 Applications in metal sensing

The rapid escalation of agricultural and industrial activities as a result of the population growth is yielding a dramatic proliferation of the amount of pollutants released daily worldwide into the environment.²⁸⁵⁻²⁸⁷ Metal ions in the aqueous environment have determined various diseases and seriously threaten ecosystem and public health with the rapid development of the industry in recent years.^{288, 289} A great effort has been made to fabricate portable sensors for monitoring heavy metals in the environment. Within this framework also 2DMs-based materials have been integrated in different types of sensors capable of detecting heavy and alkali metal ions *via* electrical^{92, 290-305} and optical outputs.^{306-317 306-317} The interaction between 2DMs and metal ions have been extensively explored in the past years and as a result outstanding adsorption capabilities have been achieved, opening new avenues in the field of wastewater purification and sensing.⁸⁹ Several excellent reviews have been published on the desalination of water using neat or functionalized GO,³¹⁸⁻³²¹ these reports discuss a variety of approaches to separate alkali metal ions from either seawater or wastewater through incorporation of GO or functionalized GO in membranes. Yet, those review articles do not discuss the use of 2DMs as platforms for sensing devices and in particular for detection of (heavy) metal ions. Many efforts have been made to develop new robust technologies for low-cost and effective portable sensors for monitoring heavy metals in the environment. Among various approaches, those based on physisorption or chemisorption, relying on the capturing of the pollutant (i.e. analyte) by an adsorbent (i.e. receptor), are chemically programmable as they exploit supramolecular recognition events. Understanding the dynamic adsorption capabilities of metal ions on the surface of 2DMs is therefore extremely important and will be discussed prior to the use of 2DMs in sensing devices for heavy metal ions.

3.1 The adsorption process

The development of *ad hoc* receptors of metal ions makes it possible to exploit the reversible processes of adsorption and desorption as extremely versatile strategies towards sensing in a variety of environments. Moreover, by relying on the reversible nature of non-covalent interactions, the sensor can exhibit a quick response, a fast recovery rate (i.e. real-time monitoring) and a facile regeneration to enable its use multiple times. 2DMs are atomically

thick and possess two planar surfaces available for metal ion adsorption, thus featuring extremely high surface-to-volume ratios. In particular, GO is very interesting for the removal of metal ions due to its unique hydrophilic nature and the presence of functional groups containing oxygen atoms, which can efficiently bind the metal ions to form strong surface complexes.^{78, 322} On the other hand, TMDs intrinsically possess numerous chalcogen atoms, which can act as a potential coordination sites for certain heavy metal ions. The adsorption capabilities of 2DMs will be evaluated through the maximum adsorption capacity (q_{\max}) defined as the ratio between the maximum loaded mass of analyte expressed in milligrams and the mass of the absorbent expressed in grams.

3.1.1 Metal sensing with graphene, graphene oxide and related composites

Graphene-based materials are being considered as the most promising absorbents for capturing various heavy metal ions.³²³⁻³²⁵ Remarkably, graphene can be easily produced in the form of GO, which displays numerous oxygen-rich functional groups such as carbonyls, epoxides, and hydroxides that act both as reactive sites for further covalent functionalization, and can interact *via* dipole-dipole or strong electrostatic interactions with the metal ions, enhancing the occurrence of adsorption event. While several excellent reviews have been published on the desalination of water using neat or functionalized GO,³¹⁸⁻³²¹ these reports discuss a variety of approaches to separate alkali metal ions from either seawater or wastewater and do not discuss the abilities of other 2DMs as platforms to adsorb (heavy) metal ions. Conventional methods for the quantification of heavy metal ions include plasma mass spectrometry (ICP-AES), atomic absorption/emission spectroscopy (e.g. F-AAS, UV-VIS), and polarography. To define the concentration of the remaining heavy metal ions after adsorption we should calculate the difference between the initial (C_0 ; mg L⁻¹) and the equilibrium (C_e ; mg L⁻¹) concentration. The equilibrium sorption capacity and time-dependent capacity were determined using Eq.2, where q_e are the equilibrium amount of heavy metal ions adsorbed per unit mass (m) of adsorbent (mg/g), V is the volume of the metal ion solution.

$$q_e = \frac{(C_0 - C_e) \times V}{m_{\text{adsorbent}}} \quad (\text{Eq. 2})$$

Among various adsorption isotherm models Freundlich and Langmuir are most commonly used for estimate the maximum adsorption capacity (q_{\max}) of metal ions on both 2D and 3D

carbon-based adsorbents.^{322, 324, 326, 327} GO and chemically modified graphene oxide (CMGO) are considered as a promising adsorbents for the removal of heavy metal ions such as Pb(II),^{78, 322, 326-356} Cu(II),^{78, 327, 335, 337, 339, 347, 349, 351, 356-366} Cr(VI),^{262, 286, 337, 355, 367-377} Cd(II),^{78, 327, 340, 349, 355} Hg(II),^{335, 349, 355, 378, 379} Ni(II),^{355, 380} Co(II),³⁸¹ Mn(II),³²⁷ Pd(II),³⁸² Sr(II),³⁸³ Au(III),³⁸² As(V),³⁸⁴ and U(VI).³⁸⁵⁻³⁸⁷ Moreover, in order to increase the adsorption capacity and simplify the separation of the GO from water, numerous hybrid materials have been investigated and include combination of GO/CMGO with poly(acrylamide) (PAM),^{328, 331} poly(*N*-vinylcarbazole) (PVK)³²⁹, hyperbranched polyamine (HPA),³³⁰ iron oxide (Fe₃O₄),^{333, 334, 338} polyamidoamine dendrimers (PAMAMs)³²⁷, ethylenediamine tetraacetic acid (EDTA),^{335, 336} chitosan (CS),^{337, 339, 350, 374, 382} *etc.* Hybrids of GO with other polymers have also been used to remove organic contaminants from water.³⁸⁸

Among numerous factors affecting the adsorption capacity, initial concentration of the solution, phase contact time, temperature and pH are found to play a crucial role in the process of adsorption of metal ions on GO.

The adsorption properties of neat GO towards divalent metal ions (copper, zinc, cadmium and lead) were investigated by Sitko *et al.*⁷⁸ In this seminal work, it was shown that GO prepared *via* the oxidation of synthetic graphite flakes using potassium dichromate possesses an impressive maximum adsorption capacity (q_{max}) for Cu(II) (223 mg g⁻¹), Zn(II) (345 mg g⁻¹), Cd(II) (530 mg g⁻¹), and Pb(II) ions (1120 mg g⁻¹). The single and competitive adsorption of Cu(II), Zn(II), Cd(II) and Pb(II) shows that the affinities of GO for these metal ions follow the order of Pb(II) > Cu(II) >> Cd(II) > Zn(II). Since then, many groups have investigated the impact of GO functionalization (with chemical groups and inorganic compounds) on its maximum adsorption capacity. The most relevant results reported so far have been listed in Table 3.

Table 3. Maximum adsorption capacities of CMGO composites used in Pb(II) removal process.

<i>Adsorbent</i>	q_{max} (mg g ⁻¹)	<i>Conditions</i>	<i>Ref.</i>
<i>Few layers GO</i>	1850 ^a	pH = 6, T = 333 K	³²²
	758 ^a	pH = 5.5, T = 333	³³²
<i>GO</i>	1119 ^a	pH = 5, T = 298 K	⁷⁸
<i>rGO/PAM</i>	1000 ^a	pH = 6, T = 298 K	³²⁸

<i>PVK-GO</i>	888 ^a	pH = 7, T = 298 K	329
<i>HPA-GO</i>	820 ^a	pH = 5.9, T = 318 K	330
<i>PAM-G</i>	820 ^a	pH = 6, T = 288 K	331
<i>MnFe₂O₄/GO</i>	673 ^a	pH = 5, T = 298 K	333
<i>GO/Fe₃O₄</i>	588 ^a	pH = 5, T = 303 K	334
<i>GO/PAMAMs</i>	568 ^b	pH = 6, T = 283 K	327
<i>EDTA/CMGO</i>	508 ^a	pH = 4.2, T = 298 K	335
<i>EDTA/GO</i>	479 ^a	pH = 6.8, T = 298 K	336
<i>CS/GO</i>	461 ^c	pH = 6, T = 318 K	337
	99 ^a	pH = 6, T = 298 K	352
<i>Fe₃O₄/Cysteine</i>	459 ^a	pH = 6, T = 298 K	338
<i>CS/GO-SH</i>	447 ^b	pH = 5, T = 293 K	339
<i>Polydopamine/GO</i>	365 ^a	pH NA, T = 298 K	340
<i>NH₂-SiO₂/GO</i>	345 ^a	pH = 5, T = 313 K	341
<i>Mesoporous silica/GO</i>	333 ^a	pH = 7.1, T = 298 K	342
<i>Ag/GO</i>	313 ^a	pH = 5.3, T = 298 K	343
<i>Hydroxyapatite/GO</i>	278 ^a	pH = 4.5, T = 308 K	344
<i>Polysiloxane/GO</i>	256 ^a	pH = 5, T = 313 K	345
<i>Phenylenediamine/ rGO</i>	228 ^a	pH = 7, T = 298 K	346
<i>Tryptophan/GO</i>	222 ^a	pH = 4, T = 293 K	347
<i>GO/Polyaniline</i>	217 ^a	pH = 5, T = 303 K	348
<i>CS/FeOOH/GO</i>	111 ^a	pH = 5.5, T = 313 K	350
<i>GO-SH</i>	108 ^a	pH = 4-10, T = 298 K	351
<i>GO-NH₂</i>	96 ^a	pH = 6, T = 298 K	326
<i>MHCGO</i>	79 ^a	pH = 5.5, T = 298 K	353
<i>CMGO</i>	77 ^a	pH = 5, T = 298 K	354
<i>Fe₃O₄-G</i>	28 ^b	pH = 6, T = 293 K	355
<i>NH₂-SiO₂/GO</i>	14 ^a	pH = 5.5, T = 298 K	356

List of abbreviations. rGO : reduced graphene oxide, PAM : poly(acrylamide), PVK : Poly(*N*-vinylcarbazole), HPA : hyperbranched polyamine, MnFe₂O₄ : manganese iron oxide, Fe₃O₄ : iron oxide, PAMAM : polyamidoamine, EDTA : Ethylenediaminetetraacetic acid, CMGO : chemically modified graphene oxide, CS : chitosan, GO-SH : Sulfhydryl-functionalized graphene oxide, NH₂-SiO₂ : amino siloxane, FeOOH : iron (III) oxide-hydroxide, GO-NH₂: aminosilanized graphene oxide, MHCGO : magnetic carboxymethyl chitosan. Q_{max} values were calculated using (a) Langmuir, (b) Freundlich or (c) Redlich Peterson isotherms.

Zhao *et al.* demonstrated that q_{max} of two- or three-layer thick GO nanosheets synthesized from flake-graphite through modified Hummers method can be tuned with temperature.³²²

Interestingly, it has been shown that the adsorption of Pb(II) ions on few-layer thick GO nanosheets was independent of the ionic strength, i.e. not affected by concentrations of background electrolyte (NaClO₄) at the pH ranging from 1 to 13. In particular, q_{\max} of Pb(II) ions calculated from the Langmuir model amount to 842, 1150, and 1850 mg g⁻¹ at 293, 313, and 333 K, respectively. Yang *et al.* proposed a different approach which relies on the functionalization of rGO with water-soluble poly(acrylamide) (PAM).³²⁸ The GO prepared following the Staudenmaier method was thermally reduced yielding rGO. PAM is a polymer with a large number of acetylamine groups in its macromolecular chains, which can interact with metal ions *via* coordination. In particular, the carboxyl groups at the periphery of rGO sheets were converted to amine groups by reaction with N-hydroxysuccinimide and 1,3-diaminopropane, and a free-radical polymerization initiator was anchored to the rGO sheets. The highest adsorption capacities of rGO/PAM for Pb(II) amounted to 1000 mg g⁻¹ (298 K), which is comparable to the neat rGO.

Aqueous solutions containing copper ions have been exploited for technological applications in the fields of mechanical manufacturing, electroplating, light industry and architecture. Yet, these solutions may cause serious diseases in the human central nervous system.³⁶¹ According to EPA regulations, copper concentration in drinking water should not exceed 1300 ng mL⁻¹.³⁵¹ Tan *et al.* prepared a hybrid composite including L-tryptophan (L-Trp) and GO by nucleophilic substitution reaction in order to increase the hydrophobicity of GO and to promote the sorption.³⁴⁷ It was found that such chemical modification of GO increases its sorption capacity from 223 to 588 mg g⁻¹. It has been shown that that high removal efficiencies can be obtained reaching values exceeding 95% at pH 5 and 4 for Cu(II) and Pb(II), respectively.

Several examples of CMGO-based adsorbents rely on the use of chitosan as molecule possessing a high affinity for heavy metal ions.^{337, 339, 350, 374, 382} Chitosan reacts with GO carboxyl groups of GO and forms amide bonds.^{262, 354} In particular, chitosan modified GO has been prepared *via* covalent modification and electrostatic self-assembly by Li *et al.*³³⁹ The introduction of GO-SH sheets as an interlayer can offer extra space into the chitosan structure and further increase the specific surface area. The results indicated that a new type of sorbent material, with functional groups such as -OH, -COOH, -SH and -NH₂ has a high adsorption capacity of copper ions (425 mg g⁻¹, see Table 4). Furthermore, the recyclability of the sorbent has been studied by treating it with HNO₃ and EDTA solutions. The chitosan (CS)/GO-SH revealed a decrease of the adsorption capacity over three cycles by 23% for Cu,

25% for Pb and 26% for Cd ions. Chitosan/GO hybrid in the form of nanofibrous composite has been studied by Najafabadi *et al.*³³⁷ The nanofibrous morphology of the hybrid material has been achieved through the use of electrospinning process. The maximum Cu(II) adsorption capacity was estimated as 423.8 mg g⁻¹. It was shown that the q_{max} decreased slowly with the increasing cycle number; such behaviour was attributed to the decrease in availability of active sites of adsorbent for metal ions. CS/GO-SH nanofibers could be used up to the fifth cycle of regeneration using HNO₃ solution by retaining 91.5% of the initial adsorption capacity Cu(II) ions sorption.

Table 4. Maximum adsorption capacities of CMGO composites used in Cu(II) removal process.

<i>Adsorbent</i>	<i>q_{max} (mg g⁻¹)</i>	<i>Conditions</i>	<i>Ref.</i>
<i>Trp/GO</i>	588 ^a	pH = 5, T = 293 K	347
	425 ^a	pH = 5, T = 293 K	339
<i>CS/GO</i>	424 ^c	pH = 6, T = 318 K	337
	203 ^b	pH = 5, T = 293 K	358
	162 ^a	pH = 5.5, T = 303 K	359
<i>Polyallylamine/GO</i>	54 ^a	pH = 5, T = 293 K	366
	349 ^a	pH = 6, T = 293 K	357
<i>EDTA/CMGO</i>	301 ^a	pH = 5, T = 298 K	335
<i>GO</i>	294 ^a	pH = 5, T = 298 K	78
	117 ^b	pH = 5.3, T = 293 K	361
<i>GO/CdS</i>	137 ^a	pH = 6, T = 298 K	360
<i>GO-EDTA</i>	109 ^a	pH = 5, T = 293 K	362
<i>GO-NH₂</i>	103 ^a	pH = 6, T = 298 K	363
<i>GO-SH</i>	100 ^a	pH = 6, T = 298 K	363
	42 ^a	pH = 5-10, T = 293 K	351
<i>GO/PAMAMs</i>	69 ^a	pH = 4.5, T = 298 K	327
<i>SMGO</i>	63 ^a	pH = 4.7, T = 323 K	364
<i>Alginate/GO</i>	60 ^a	pH = NA, T = 293 K	365
<i>NH₂-SiO₂/GO</i>	6 ^a	pH = 5.5, T = 293 K	356

List of abbreviations. Trp : tryptophan, CS : chitosan, EDTA : Ethylenediaminetetraacetic acid, CMGO : chemically modified graphene oxide, CdS : cadmium sulfide, GO-NH₂ : aminosilanized graphene oxide, GO-SH : Sulfhydryl-functionalized graphene oxide PAMAM : polyamidoamine, SMGO : sulfonated magnetic graphene oxide, NH₂-SiO₂ : amino siloxane.

Q_{max} values were calculated using (a) Langmuir, (b) Freundlich or (c) Redlich Peterson isotherms.

Chromium, which exists most frequently as Cr(III) and Cr(VI),³⁷⁷ is commonly used in metallurgy, tanning, military, dyes and pigments.^{368, 369} Generally, Cr(VI) compounds are considered to be more toxic than Cr(III) because of their toxicity to humans, animals, plants and microorganisms.^{377, 389} A variety of chemical approaches has been employed in order to increase the affinity of GO towards chromium ions. For example, functionalization of GO with polypyrrole (PPy), being a widely studied conductive polymer, has been exploited by Li *et al.*³⁶⁸ The PPy/GO composite nanosheets prepared by sacrificial template polymerization method exhibit adsorption capacity for Cr(VI) ions as high as 497 mg g⁻¹. Moreover, the adsorption capacity of the PPy/GO composite nanosheets is about twice as large as that of conventional PPy nanoparticles.^{390, 391} Supplementary studies showed that subsequent covalent functionalization of PPy/GO with α -cyclodextrin³⁶⁷ leads to maximum adsorption capacities of 606 mg g⁻¹ at 25 °C and 667 mg g⁻¹ at 45 °C.

Table 5. Maximum adsorption capacities of CMGO composites used in Cr(VI) removal process.

<i>Adsorbent</i>	<i>q_{max} (mg g⁻¹)</i>	<i>Conditions</i>	<i>Ref.</i>
<i>PPy/cyclodextrin/GO</i>	667 ^a	pH = 2, T = 318 K	367
<i>PPy/GO</i>	497 ^a	pH = 3, T = 293 K	368
	310 ^c	pH = 3, T = 318 K	337
<i>CS/GO</i>	219 ^a	pH = 2, T = 293 K	371
	145 ^a	pH = 3-4, T = 303 K	374
	108 ^a	pH = 2, T = 303 K	376
<i>Fe₃O₄/GO</i>	259 ^b	pH = 2, T = 298 K	369
<i>Trioctylamine/GO</i>	232 ^a	pH = 2.5-3, T = 293 K	370
<i>G-MgAl-LDH</i>	172 ^b	pH = 2, T = 293 K	372
<i>Fe/G</i>	162 ^a	pH = 4.5, T = 293 K	373
<i>Cyclodextrin/GO</i>	120 ^a	pH = 3, T = 293 K	286
<i>Fe₃O₄/TiO₂/GO</i>	118 ^b	pH = 2, T = 303 K	375
<i>DCTA/GO</i>	84 ^b	pH = 2, T = 303 K	377
<i>Fe₃O₄/G</i>	17 ^b	pH = 1-3, T = 293 K	355

List of abbreviations. PPy : polypyrrole, CS : chitosan, Fe₃O₄ : iron oxide, TiO₂ : titanium dioxide, G-MgAl-LDH : graphene - magnesium/aluminum - layered double hydroxide, DCTA : 1,2-diamino cyclohexanetetraacetic acid.

Q_{max} values were calculated using (a) Langmuir, (b) Freundlich or (c) Redlich Peterson isotherms.

Numerous organic and inorganic compounds have been used to functionalize GO and form composites in order to detect and capture other heavy metal ions like Cd(II),^{78, 327, 340, 349, 355} Hg(II),^{335, 349, 378, 379} Ni(II),^{355, 380} Co(II),³⁸¹ Mn(II),³²⁷ Pd(II),³⁸² Sr(II),³⁸³ Au(III),³⁸² As(II)³⁸⁴ and U(III)³⁸⁵⁻³⁸⁷ efficiently from aqueous solutions. Zhang *et al.* have demonstrated that efficient CMGO-based sorbent can be achieved by decorating GO with polyamidoamine dendrimers (GO/PAMAMs).³²⁷ The adsorption behaviour of GO/PAMAMs for heavy metal ions in water solution was studied by changing the concentration of heavy metal ions, pH values, and temperature. The maximum adsorption capacities of GO/PAMAMs were found to be 568.18, 253.81, 68.68, and 18.29 mg g⁻¹ for Pb(II), Cd(II), Cu(II), and Mn(II), respectively. Noteworthy, it was also found that adsorption capacities of the GO/PAMAMs for the heavy metal ions were highly pH dependent. In particular, at pH < 3, the hydronium ions of higher concentration compete with M(II) to grasp the adsorption sites. With an increase of pH, the protonation degree of the amino groups was weakened, and the coordination and chelating ability of PAMAM's amino groups toward Pb(II), Cd(II), Cu(II), and Mn(II) reinforced.

Conducting polymers interfaced with carbon and carbon-based derivatives have displayed enhanced removal of mercury and other toxic materials from water. A facile chemical route was reported by Chandra *et al.*³⁷⁸ In particular, they showed that polypyrrole (PPy)-rGO composites possess a highly selective Hg²⁺ removal capacity. rGO sheets cross-linked with polypyrrole exhibited an increased surface area of 166 m² g⁻¹. The uptake of Hg²⁺ by PPy/rGO has been estimated being as high as 980 mg g⁻¹. Furthermore, PPy/rGO possess an extremely high desorption capacity of up to 92.3%. Chitosan/GO (CSGO) composites with three different loadings of GO, i.e. 5, 10 and 15 wt% were prepared for the adsorption of Au(III) and Pd(II) by Liu *et al.*³⁸² The adsorption capacity of Au(III) and Pd(II) onto CSGO composites revealed high values at pH 3.0–5.0 for Au(III) and pH 3.0–4.0 for Pd(II). It was found, that the composite with 5 wt% of GO had the largest adsorption capacity for Au(III) and Pd(II) compared with the other prepared adsorbents, where the maximum adsorption capacity amounted to 1076.6 mg g⁻¹ and 216.9 mg g⁻¹ for Au(III) and Pd(II), respectively.

Noteworthy, GO-based composites can also be used also in adsorption process of radioactive ions.³⁸⁵⁻³⁸⁷ Chen *et al.* reported on the amino functionalized magnetic graphene oxide composite (AMGO) synthesized by a facile, one-step solvothermal method, tailor made for the U(VI) removal from aqueous solutions.³⁸⁷ It was shown, that the sorption of U(VI) on

AMGO occurs *via* the formation of coordination complexes with the nitrogen- and oxygen-containing functional groups. It was concluded that the chemical affinity of the U(VI) for the nitrogen containing functional groups is stronger than that for the oxygen containing functional groups. Interestingly, the AMGO composite could be recovered from the solution with the magnetic separation within one minute. The same group has demonstrated that the adsorption capacity of CMGO towards uranium can be improved by functionalizing GO with activated carbon felt (ACF) through electrophoretic deposition and subsequent thermal annealing.³⁸⁵ The q_{\max} of GO-ACF for U(VI) amounted to 298 mg g⁻¹ at pH 5.5, being much higher than the one of ACF (173 mg g⁻¹), thus suggesting that the carboxyl functional groups of GO-ACF play a salient role in the sorption process, yielding a high efficiency for the removal of U(VI).

Table 6. Maximum adsorption capacities of CMGO composites used in various metal ions removal process.

<i>Adsorbent</i>	<i>Contaminant</i>	<i>q_{max} (mg g⁻¹)</i>	<i>Conditions</i>	<i>Ref.</i>
GO		530 ^a	pH = 5, T = 298 K	78
GO/PAMAMs	Cd(II)	253 ^b	pH = 5, T = 298 K	327
PDA/GO		210 ^a	pH = NA, T = 298 K	340
Fe ₃ O ₄ /G		28 ^b	pH = 6-7, T = 293 K	355
PPy-rGO		980 ^a	pH = 3, T = 293 K	378
Ferrite/CS/G	Hg(II)	361 ^a	pH = 7, T = 323 K	379
EDTA-GO		268 ^b	pH = 4.1, T = 298 K	335
Fe ₃ O ₄ -G		23 ^b	pH = 6-7, T = 293 K	355
GO-G	Ni(II)	37 ^a	pH = 6, T = 293 K	380
Fe ₃ O ₄ -G		22 ^b	pH = 6-7, T = 293K	355
GO-NH ₂	Co(II)	116 ^a	pH = 6, T = 298 K	381
GO/PAMAMs	Mn(II)	18 ^a	pH = 4, T = 298 K	327
CS/GO	Pd(II)	216 ^a	pH = 3, T = 323 K	382
PAM/GO	Sr(II)	185 ^a	pH = 8.5, T = 303 K	383
CS/GO	Au(III)	1076 ^a	pH = 4, T = 303 K	382
GO-FeOOH	As(V)	73 ^a	pH = 7, T = 298 K	384
GO/AC		298 ^a	pH = 5, T = 298 K	385
GO-sepiolite	U(VI)	161 ^a	pH = 5, T = 298 K	386
GO-NH ₂		141 ^a	pH = 6, T = 298 K	387

List of abbreviations. PAMAM : polyamidoamine, PDA: polydopamine, Fe₃O₄: iron oxide, PPy : polypyrrole, CS: chitosan, EDTA: Ethylenediaminetetraacetic acid, GO-NH₂ : aminosilanized graphene oxide, PAM : poly(acrylamide), FeOOH : iron (III) oxide-hydroxide, AC : activated carbon.

Q_{max} values were calculated using (a) Langmuir or (b) Freundlich isotherms.

3.1.2 Metal sensing with transition metal dichalcogenides

TMDs are characterized by an abundance of intrinsic chalcogen atoms, which are suitable coordination sites for certain heavy metal ions. Because of this reason, TMDs are now being considered as promising sorbents. In particular, due to the strong non-covalent sulphur-mercury and sulphur-lead interactions MoS₂ has been recently investigated for the efficient purification of wastewater. The thickness of the S-Mo-S triple layer in MoS₂ amounts to 3.17 Å, resulting in the distance of 2.98 Å between two neighbouring layers.¹⁴⁰ Such a distance is too narrow to allow the penetration of hydrated chalcogen ions in the interior spaces where the vast majority of potential coordination sites are located.³⁹² Therefore, maximizing the number of accessible sulphur atoms is a crucial step for using layered TMDs in the high-efficiency removal of heavy metal ions from contaminated water. Numerous papers have been published on the control over the interlayer spacing of MoS₂, either by post-treatment of bulk MoS₂ or *in situ* synthesis. For example, Gao *et al.* reported the preparation of MoS₂ with interlayer spacing of 9.4 Å *via* the intercalation of oxidized DMF species.³⁹³ Zheng *et al.* demonstrated that the interlayer spacing of MoS₂ could be increased to 7.08 Å, 8.99 Å, and 11.61 Å by intercalation with Li⁺, K⁺ and Na⁺, respectively.³⁹⁴ Recently, Ai *et al.* showed that the interlayer spacing of MoS₂ can also be controlled by the synthesis temperatures.³⁹⁵ In particular, the interlayer spacing of the obtained MoS₂ was greatly enlarged (9.5 Å) in the temperature range of 140–200 °C; upon further increase of the temperature up to 220 °C, the interlayer spacing of the obtained MoS₂ was found to be almost identical to that of bulk MoS₂. Additionally, extremely fast adsorption kinetics during efficiently reduced the mercury concentration have been observed (Figure 11a). In brief, 20 mg widened defect-rich nano MoS₂ (W-DR-N-MoS₂) was added to 200 mL of aqueous solution containing 10 000 ppb Hg²⁺. The concentration of Hg²⁺ drastically decreased to 16.9 ppb after 5 min of treatment, indicating a removal efficiency of 99.83%. The structural features were systematically investigated of W-DR-N-MoS₂ (Figure 11b-e). They revealed for the first time that such widened defect-rich MoS₂ (W-DR-N-MoS₂) nanosheets are capable to capture Hg(II) ions, with an extremely high q_{max} (2563 mg g⁻¹) (Figure 11f) closely matching the theoretically

predicted value of 2506 mg g^{-1} (assuming a stoichiometric S/Hg ratio of 1:1), being higher than the q_{max} of the best adsorbents to date.³⁹⁶⁻³⁹⁸ Selectivity of W-DR-N-MoS₂ exhibited negligible capturing capability for various competitive ions, such as Na⁺, K⁺, Ca²⁺, Mg²⁺, Al³⁺, Cr³⁺, Mn²⁺, Zn²⁺ and Cd²⁺, and even inferior capturing capacity for Fe³⁺, Cu²⁺ and Pb²⁺.

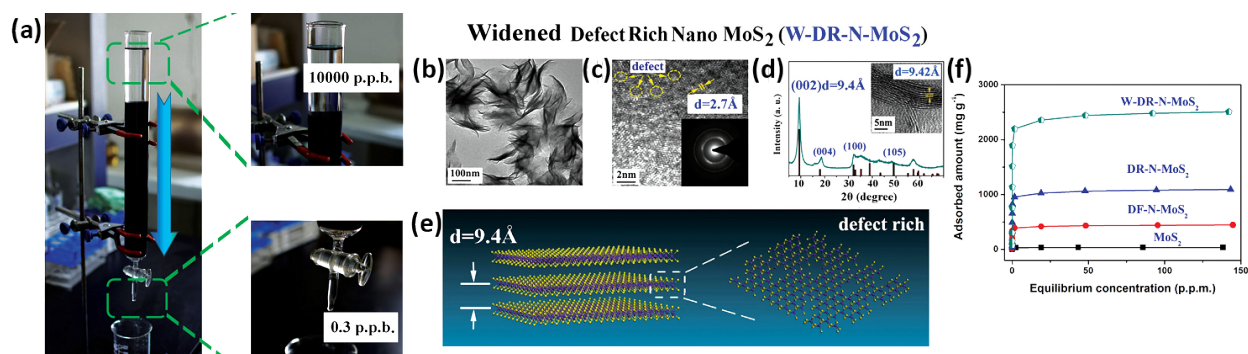


Figure 11. (a) Purification of natural water sample. The mercury-contaminated water was passed through a purification column filled with W-DR-N-MoS₂: (b) TEM image of W-DR-N-MoS₂. (c) HRTEM image of the basal plane of W-DR-N-MoS₂, inset is the corresponding SAED pattern. (d) XRD pattern and the cross-sectional HRTEM image of W-DR-N-MoS₂. (e) Structural model of W-DR-N-MoS₂ with enlarged interlayer spacing and multiple defects on the basal planes (The sulphur atoms are yellow and the Mo atoms are purple). (f) Mercury adsorption isotherm, compared with commercial MoS₂ powder, DF-N-MoS₂ and DR-N-MoS₂. Adapted from Ref.³⁹⁵ with permission of Wiley-VCH.

Liu *et al.* showed that MoS₂ exhibits superior ability to adsorb lead ions.³⁹⁹ MoS₂ nanosheets were prepared with an ultrasound assisted electrochemical exfoliation method. As in the case of Hg(II) ions the adsorption capacity of MoS₂ towards Pb(II) can be attributed to the coordination of Pb(II) with S atoms of MoS₂. Moreover, the strength of such non-covalent interactions is being reinforced by electrostatic adsorption. The experimental Pb(II) uptake capacity of MoS₂ was estimated being as high as 1479 mg g^{-1} . The adsorption followed the Freundlich isotherm model and fitted well with both the pseudo-first-order and pseudo-second-order kinetics models. The adsorption of Pb(II) on MoS₂ materials was further confirmed by the measurement of SEM-EDS (Figure 12a-d) and XPS measurements (Figure 12e). The immobilization of heavy metal ions on 2DMs is depicted on Figure 12f. The adsorption might be attributed to the chemical adsorption due to the complexation of Pb(II) with intrinsic S or O atoms exposed on MoS₂ surfaces, together with electrostatic adsorption.

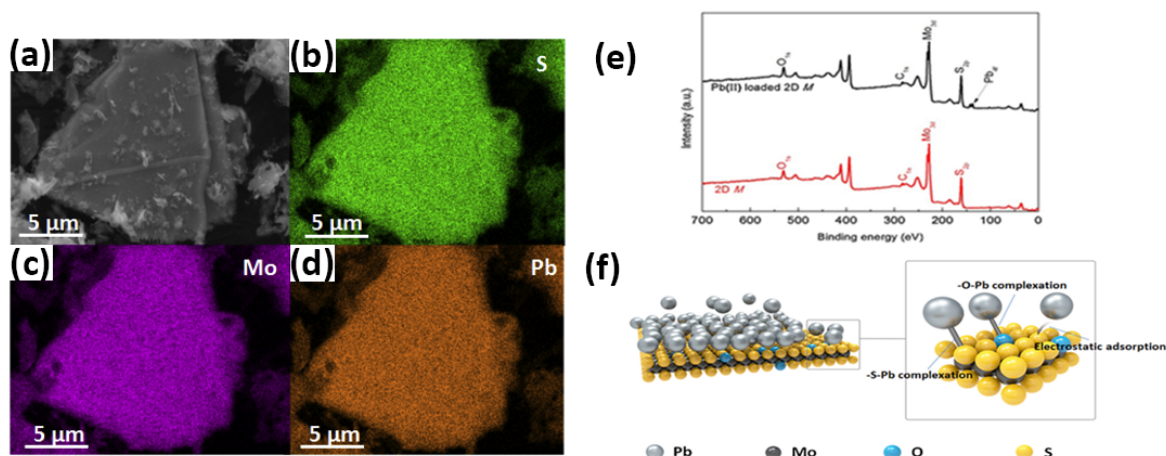


Figure 12. (a) SEM image of Pb(II) loaded 2DM. b–d) SEM-EDS elemental mapping images of S, Mo, and Pb, respectively. (e) Diagrammatic illustration of the mechanism for Pb(II) adsorption on 2DM. (f) XPS survey spectra of 2DM before and after Pb(II) adsorption. Adapted from Ref. ³⁹⁹ with permission of Elsevier Ltd.

Other studies showed that the modification or the use of MoS₂ as a component of a hybrid material does not increase the q_{\max} towards Pb(II)^{399, 400} and Hg(II),^{395, 401, 402} (see Table 7).

Table 7. Maximum adsorption capacities of TMDs composites used in Hg²⁺ and Pb²⁺ metal ions removal process.

<i>Adsorbent</i>	<i>Contaminant</i>	q_{\max} (mg g ⁻¹)	<i>Conditions</i>	<i>Ref.</i>
<i>W-DR-N-MoS₂</i>	Hg(II)	2563 ^a	pH = NA, T = NA	395
<i>Au/Fe₃O₄/MoS₂</i>		1527 ^a	pH = 5, T = 293 K	402
<i>MoS₂</i>		305 ^b	pH = 6, T = 308 K	401
<i>MoS₂</i>	Pb(II)	1479 ^b	pH = 5, T = 308 K	399
<i>Co-MoS₂</i>		660 ^a	pH = 1-6, T = 293 K	400
<i>Mn-MoS₂</i>		588 ^a	pH = 1-6, T = 293 K	400

List of abbreviations. W-DR-N-MoS₂: widened defect rich nano molybdenum disulfide, Au: gold.

Q_{\max} values were calculated using (a) Langmuir or (b) Freundlich isotherms.

3.2 Fluorescence-based metal sensors

Colorimetric sensors for metal ions comprise two key features, i.e. a metal chelating or binding (coordination) pocket and at least one fluorophore capable of absorbing and/or emitting light. Fluorescence sensing is based on analyte-induced changes in the physicochemical properties of fluorophores including fluorescence intensity, lifetime, and anisotropy, which are related to charge transfer or energy transfer processes.⁴⁰³ To function as a sensor, the electronic structure of the sensor must be altered upon metal binding. Changes in the electronic structure of the sensor can lead to the changes in the intensity or wavelength of light absorption or emission, while changes in the molecular structure can modify the distance or alignment between a pair of fluorophores that serve as a donor–acceptor pair.

3.2.1 Graphene-based fluorescence metal sensors

Graphene oxide has superb fluorescence quenching capability.^{404, 405} For example, a fluorescent quantum dots (QD)–aptamer–GO sensor based on the nanometal surface energy transfer (NSET) mechanism was designed by Qian *et al* in order to detect Pb^{2+} ions.³⁰⁶ In particular, photoinduced electron transfer between graphene quantum dots and GO was employed to achieve the controllable fluorescence turn-on process. The capability of QD–aptamer–GO nanosensor to detect Pb(II) and other ions with high sensitivity and good reproducibility was demonstrated on Figure 13.

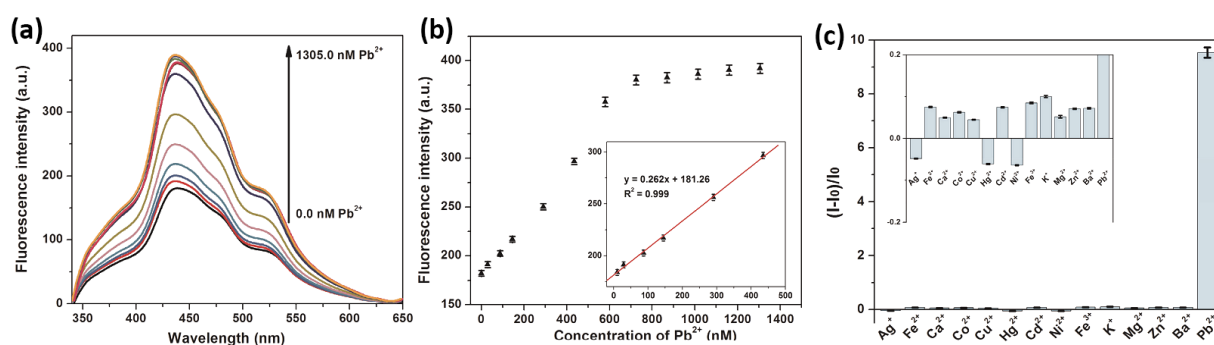


Figure 13. (a) Fluorescence recovery of aptamer–rGQDs/GO system after incubation with various concentrations of Pb^{2+} (0.0, 29.0, 87.0, 145.0, 290.0, 435.0, 580.0, 725.0, 870.0, 1015.0, 1160.0, 1305.0 nM). (b) Linear relationship between the fluorescence intensity and concentration of Pb^{2+} . (c) Fluorescence intensity changes ($(I-I_0)/I_0$) of the sensor in the presence of various metal ions. Adapted from Ref. ³⁰⁶ with permission of Elsevier Ltd.

Li *et al.* demonstrated that the single-stranded DNA (ssDNA) aptamer attached on GO can specifically bind to the mercury ions, leading to the formation of a hairpin-shaped double-stranded DNA (dsDNA) structure.³⁰⁷ The water-dispersible GO sheets, which are functionalized with ssDNA aptamer, exhibit strong fluorescence emission at 600 nm under the excitation of 488 nm in the absence of Hg(II) ions. When Hg²⁺ ions appear in the aqueous solution, they are being sandwiched between the hairpin-shaped dsDNA due to the formation of the thymine–Hg(II)–thymine complex, which grasps the Hg²⁺ ions in proximity to the surface of GO. As a result, the fluorescence emission of GO is quenched. Such sensor shows a limit of detection as low as 0.92 nM and excellent selectivity towards Hg(II) over a wide range of metal ions including K⁺, Ag⁺, Ca²⁺, Cd²⁺, Cu²⁺, Pb²⁺, Ni²⁺, Co²⁺ and Fe³⁺. Recently, Wen *et al.* reported a FRET sensor based on a cytosine rich DNA probe and GO.³⁰⁸ In the presence of Ag⁺, a DNA–Ag⁺ complex was formed, and the conformation of the probe changed to straight stiff, resulting in desorption of DNA from the surface of GO and fluorescence recovery (Figure 14a). Fluorescence spectra of cytosine-rich oligonucleotide (SSO) probe upon incubation with a series of concentrations of Ag⁺ and then mixed with GO is portrayed in Figure 14b. This assay is based on the interaction between the target-induced conformational change of the SSO fluorogenic probe and graphene oxide quenching effects. The use of a simple mix-and-detect analysis revealed a high selectivity toward Ag⁺ as determined in the presence of a ten times higher concentration of 12 different interference metal ions (Figure 14c).

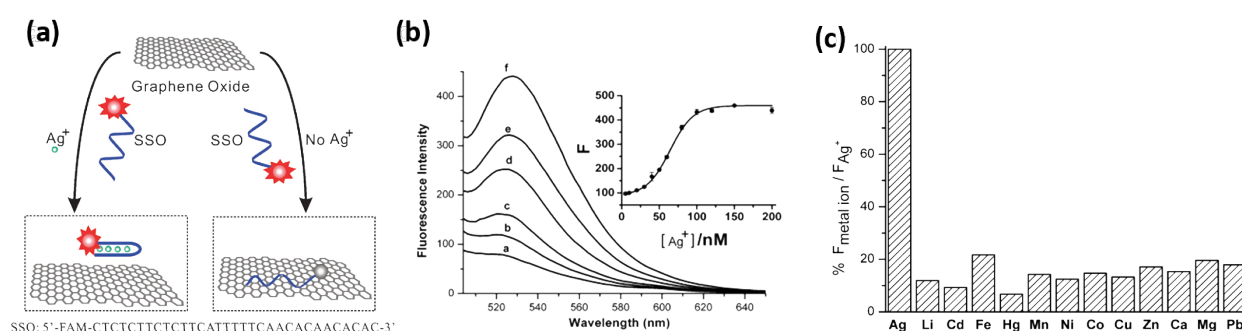


Figure 14. Schematic illustration of the fluorescence sensor for Ag(I) ions based on the target-induced conformational change of a silver specific cytosine-rich oligonucleotide (SSO) and the interactions between the fluorogenic SSO probe and graphene oxide. SSO: (a) FAM-labelled silver-specific oligonucleotide probe. (b) Fluorescence spectra of SSO probe upon incubation with a series of concentrations of Ag⁺ and then mixed with GO: *a* 0 M, *b* 20 nM, *c*

40 nM, *d* 60 nM, *e* 80 nM and *f* 150 nM. All experiments were carried out in MOPS buffer (10 mM, pH 7.0) containing 50 mM of NaNO₃ and 10 nM of SSO. Ag⁺ of different concentrations were incubated in this solution for 5 min at 23°C and then the fluorescence spectra were recorded 2 min after GO (10 mg) were added to this mixture. λ_{ex} = 494 nm. (c) Selectivity of the analysis of Ag⁺ ions in the presence of different metal ions. The concentration of Ag⁺ was 100 nM whereas all other interference metal ions were 1 μM. Adapted from Ref. ³⁰⁸ with permission of the Royal Society of Chemistry.

3.2.2 Transition metal dichalcogenides based fluorescence metal sensors

As a result of their finite band gap, TMDs exhibit interesting optical properties. By mastering the selective recognition capabilities of TMDs toward specific heavy metal ions, several fluorescence sensors for metal ions detection have been developed recently, highlighting the promise of using TMDs-based materials to construct novel nanoprobe for chemical sensing and in particular for detection of metal ions.³⁰⁹⁻³¹⁴ For heavy metal ion sensors, TMDs are either non-covalently functionalized with molecules, which are used as sensing probes - because of their high binding affinity to heavy metal ions,³¹⁴ or covalently modified, e.g. through N- or B- doping.³¹¹ Mao *et al.* reported a seminal work on the use of a single layer MoS₂ as the fluorescence quencher to design a probe for detection Ag^I with excellent sensitivity and selectivity. The detection limit in this assay was found being as low as 1 nM for Ag⁺.³⁰⁹

Recently, a fluorescent nanoprobe based on MoS₂ nanosheets for selective and ultrasensitive detection of Ag⁺ (down to 10nM) ions alike in aqueous solutions and biological cells (*Escherichia coli*) has been developed by Yang *et al.*³¹⁰ In this study Ag^I ions were reduced to Ag⁰, which led to the detachment of Rhodamine B isothiocyanate (RhoBS) non-covalently interacting with MoS₂ nanosheets, which resulted in quenching of its fluorescence (see Table 8 for details).

An interesting approach for enhancing the fluorescence of MoS₂ has been proposed by Liu *et al.*³¹¹ and relies on covalent modification of MoS₂ nanosheets with either boron or nitrogen atoms. Boron and nitrogen doping results in changes of the band gap of MoS₂ which increases from 1.20 eV to 1.61 eV. The as-prepared B- and N-doped MoS₂ nanosheets were used as facile, green, label-free and effective sensing platforms for Hg²⁺ ions. The modified

nanosheets exhibit enhanced fluorescence properties compared with the undoped MoS₂, which were highly quenched after selective absorption of Hg²⁺. Wang *et al.* found out that after the doping with Pb²⁺ ions, the fluorescence of MoS₂ nanosheets was enhanced, which was subsequently quenched by the addition of sulphide ions.³¹² The fluorescence quantum yield measurements of the single layer MoS₂ was estimated being ca. 0.28%. The quantum yield of the doped MoS₂ nanosheets was calculated to be ca. 0.73%, thereby indicating that molecular doping of MoS₂ nanosheets with lead(II) metal ions strengthened their fluorescence.

To date little has been done on the application of WS₂ nanosheets in biological and chemical sensing. For example, Zuo *et al.* developed a novel dual-colour fluorescent biosensing platform based on WS₂ nanosheets.³¹³ The sensor could achieve simultaneous detection of Hg²⁺ and Ag⁺ in a high sensitivity and wide linear range by monitoring fluorescence intensity changes at 525 nm and 583 nm, respectively. Hg²⁺ and Ag⁺ were selectively detected in the concentration range from 6.0 to 650.0 nM and from 5.0 to 1000.0 nM, with detection limits of 3.3 nM and 1.2 nM, respectively. The fluorescence intensities in the presence of other heavy metal ions changed by ca. 5% when cations concentrations were 10-fold greater than that of Hg²⁺ and Ag⁺ ions, indicating that the developed biosensor is highly selective for mercury and lead ions.

3.2.3 Other 2DMs based fluorescence metal sensors

Liu *et al.* showed an integrated black phosphorus-tilted fibre grating (TFG) configuration as an enhanced light-matter interaction platform for heavy metal sensing.³¹⁵ BP nanosheets were produced by liquid-phase exfoliation and deposited through *in situ* layer-by-layer technique. The BP-TFG composite was exploited as the optic probe for Pb²⁺ ions detection, and the results showed ultrahigh sensitivity up to 0.5×10^{-3} dB/ppb, limit of detection down to 0.25 ppb, and extremely wide concentration range from 0.1 ppb to 1.5×10^7 ppb. The achievable limit of detection was estimated as 40 times lower than World health organization permissible limit for lead in drinking water. The wide detection range is 4-orders of magnitude larger than that of BP-FET based lead sensor.³¹⁶

Gu *et al.* presented novel ratiometric fluorescence sensor based on the inner filter effect (IFE) of tetraphenylporphyrin tetrasulfonic acid (TPPS) toward black phosphorus quantum dots

(BP QDs), developed for the selective and sensitive detection of Hg^{2+} .³¹⁷ Highly fluorescent BP QDs were successfully synthesized from bulk BP by sonication-assisted solvothermal method. In the presence of Hg^{2+} , the IFE originating from the spectral overlap between the excitation of BP QDs and the absorption of TPPS is inhibited; yielding in the recovery of the fluorescence of BP QDs. The constructed sensor revealed a good linear response to Hg^{2+} ranging from 1 to 60 nM with a detection limit of 0.39 nM. Interestingly, this strategy could be also applied in the determination of Hg^{2+} in real water samples with satisfactory results. Sensing properties for all 2DMs-based fluorescent sensors are presented in Table 8.

Table 8. Comparison of sensing properties of different TMDs based fluorescent sensor.

<i>Material</i>	<i>Analyte</i>	<i>Linear range</i>	<i>Detection limit</i>	<i>Ref.</i>
<i>MoS₂</i>	Ag(I)	10-500 nM	1nM	309
<i>WS₂</i>		5.0-1000 nM	1.2 nM	313
	<i>B,N-MoS₂</i>	6.0-650 nM	3.3 nM	313
Hg(II)		0.01-3 μM	1 nM	311
		1-60 nM	0.39 nM	317
<i>BP-TPPS</i>	Pb(II)	0.5-12 μM	0.22 μM	312
<i>MoS₂</i>		0.1-1.5 x 10 ⁷ ppb	0.25 ppb	315
<i>BP-fibres</i>				

List of abbreviations. TPPS : tetraphenylporphyrin tetrasulfonic acid.

3.3 Field-Effect Transistor based metal sensor

2DMs nanosheets integrated in FET have recently revealed their enormous potential for detection of heavy metals. The working principle of 2DM-based FET sensor is based on the changes of the critical parameters of a FET containing 2DMs nanosheets upon adsorption of targeted heavy metal ions. This includes primarily the field-effect mobility, threshold voltage and $I_{\text{on}}/I_{\text{off}}$ ratio. 2D semiconducting sheets are of particular interest because their high charge carrier mobility and very high surface-to-volume ratio, leading to high sensitivity. Chemical sensors based on FETs can overcome the obstacles of previous detection methods. For example, aforementioned optical methods have some limitations such as multiple sensing steps, the need for using chemical agents, a higher cost, and a longer detection time. In contrast, the use of 2DM-based FET sensors enables the rapid label-free detection of metal ions in real-time by monitoring the resistance or the Dirac point shift caused by the

adsorption of target analytes. Such devices can be characterized also by low power consumption and can be miniaturized for the development of portable sensors, eventually supported on flexible foils.

3.3.1 Graphene-based FET metal sensors

In a typical graphene field-effect transistor (GFET) based sensor, graphene is used as conducting material in the channel between drain and source electrodes. Gate potential is applied through back-gate (typical thin SiO₂ layer)⁴⁰⁶ or top-gate (electric double layer in electrolyte).¹¹⁴ The absorption of analyte molecules or change of local environment leads to the change of graphene electrical conductance. Zhang *et al.* reported the functionalization of mechanically exfoliated graphene with a self-assembled monolayer of 1-octadecanethiol and its application in Hg²⁺ sensing.²⁹⁰ Substituted alkane derivatives were found to self-assemble into large-scale highly-ordered physisorbed monolayers on single-layer graphene supported by the SiO₂ dielectric substrate. According to AFM imaging, the height of graphene increased to 1.6 nm when exposed to Hg²⁺, indicating the successful uptake of mercury ions by the thiol groups exposed on the graphene surface. Such graphene-FET sensor revealed a detection limit for Hg²⁺ as low as 10 ppm.

Alternative approaches based on solution processable rGO,^{291, 292} and G functionalization^{92, 293-295} are being pursued with the ultimate goal of developing low-cost, scalable fabrication of graphene-FET sensors. Recently, Sudibya *et al.* presented a FET sensor using micropatterned, metallothionein type II protein (MT II)-functionalized rGO films which bind with both physiological (e.g., Cu²⁺, Zn²⁺) and xenobiotic (e.g., Hg²⁺, Cd²⁺) metals ions with high affinity (see schematic illustration in Figure 15a).²⁹¹ A typical plot of drain-to-source current (I_{ds}) versus solution-gate voltage (V_g) of rGO-FET sensor is displayed in Figure 15b. Such a nanoelectronic sensor is capable of detecting various metal ions in real-time with high sensitivity. The addition of mercury (Hg²⁺), at a concentration as low as 1 nM, caused the obvious current increase in the rGO-FET which was biased at $V_{ds}= 400$ mV and $V_g=-0.6$ V (Figure 15c). The magnitude of the device response scales with the Hg²⁺ concentration, and its polarity depends on the gate voltage (V_g). The detection limit for Hg²⁺ was estimated being as low as ~1 nM with a signal-to-noise ratio of 25-30 (Figure 15d). The very same device exhibited also Cd²⁺ detection at 1 nM with a slightly smaller change of current signal-to-noise ratio of 15-20 (Figure 15e).

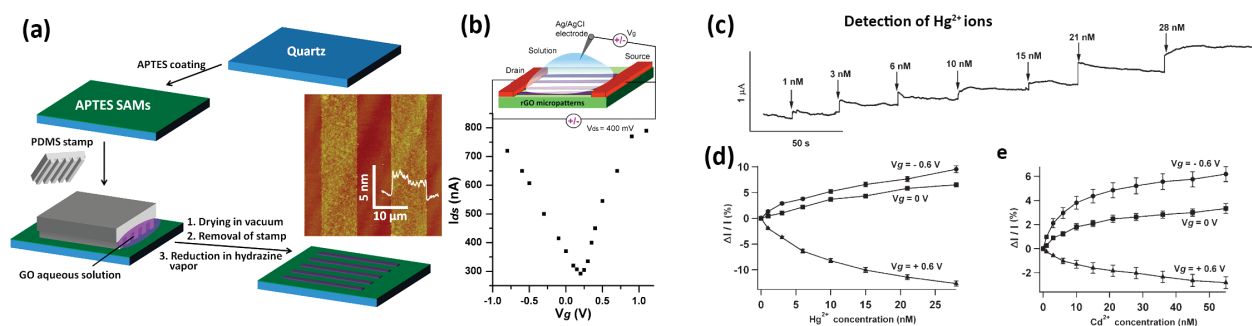


Figure 15: (a) Schematic illustration for fabrication of patterned rGO thin films on APTES-coated quartz. Inset: AFM image of the obtained rGO micropatterns on APTES-coated quartz. (b) Ambipolar characteristics of rGO-FET measured in 0.1 M of phosphate buffer saline (PBS) solution. Inset: Schematic of solution-gated configuration of rGO-FET. (c) Typical real-time recording of I_{ds} with the addition of Hg^{2+} ions. (d,e) Change of I_{ds} in rGO-FETs (tested sample number $n= 6$) with the addition of (d) Hg^{2+} and (e) Cd^{2+} ions at $V_{ds}= 0.4$ V and $V_g= -0.6$ (circle), 0 (square), and $+0.6$ V (triangle). Adapted from Ref. ²⁹¹ with permission of American Chemical Society.

For heavy metal ion GFET sensors, biomolecules are frequently used as sensing nanoprobe due to their high binding affinity to inorganic contaminants.⁹² An *et. al.*, reported high-performance flexible graphene aptasensor for Hg^{2+} detection.²⁹³ 1,5-diaminonaphthalene (DAN) and glutaraldehyde (GA) were employed as cross-linking agents, while the aptamer (3'-amine-TTC TTT CTT CCC CTT GTT TGT-C10 carboxylic acid-5') was non-covalently linked onto the graphene surface as a probe for Hg^{2+} . The field-induced responses from the graphene aptasensor had excellent sensing performance: Hg^{2+} ions with very low concentration of 10 pM could be detected (Figure 16a) being 2-3 orders of magnitude higher than previously reported mercury sensors using electrochemical devices.^{406, 407} GFET-sensor was characterized with the experimental setup presented schematically in Figure 16b. The sensor response time was rapid, with values below 1 s. Noteworthy, the pristine graphene devices showed no significant current changes upon the exposure to Hg^{2+} . It was concluded, that the origin of current changes in the sensor relies on the *p*-doping effect resulting from the thymine–Hg–thymine complex formed between Hg^{2+} ions and thymine base pairs in the aptamer (Figure 16c).

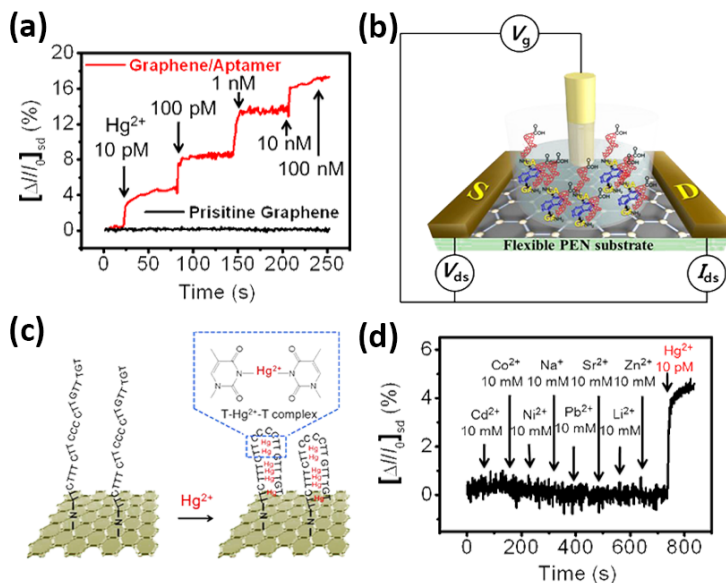


Figure 16. (a) Real-time responses curve of the aptasensor with various Hg^{2+} concentrations (10 pM to 100 nM). Graphene substrate without aptamer was introduced as a control sample. (b) Schematic diagram of a liquid-ion gated FET using graphene conjugated with aptamer. (V_g , S and D indicate gating voltage and source/drain electrodes). (c) Interaction of Hg^{2+} ions with thymine base pairs in the aptamer immobilized on the surface of the modified graphene layer (d) Selective responses of the aptasensor toward target metal ion (Hg^{2+} , 10 pM) and nontarget metal ions (Cd^{2+} , Co^{2+} , Ni^{2+} , Na^+ , Pb^{2+} , Sr^{2+} and Zn^{2+} , 10 mM). Adapted from Ref. ²⁹³ with permission of American Chemical Society.

Noteworthy, graphene-based FET sensors can be exploited not only for the detection heavy metal ions (e.g. Hg^{2+} and Pb^{2+}) but also alkali metal ions in water including K^{+294} and Na^+ .²⁹⁵

3.3.2 Transition metal dichalcogenides FET metal sensors

The semiconducting properties of MoS_2 make this 2DM particularly suitable as electroactive material for applications in transistor.⁴⁰⁸ Also because of this reason, MoS_2 has been employed in FET sensors for detecting heavy metal ions.^{296, 297}

Jiang *et al.* demonstrated the FET based on mechanically exfoliated few-layer MoS_2 nanosheets for sensing Hg^{2+} ions.²⁹⁶ The interaction between Hg^{2+} ions and few-layer MoS_2 were studied by FET measurements and photoluminescence. Due to a high binding affinity

between the sulphur sites on the surface of MoS₂ surface and Hg²⁺ ions, the latter can strongly bind to MoS₂. Remarkably, it was shown that the binding of Hg²⁺ results in a *p-type* doping and reduces the electrons concentration in *n-type* few-layer MoS₂. Upon binding of Hg²⁺ ions the electron transport and photoluminescence properties in few-layer MoS₂ are effectively modulated. It was also demonstrated, that by monitoring the changes in conductance of few-layer MoS₂ and varying the concentration of Hg²⁺ solutions, few-layer MoS₂ transistors can function as highly sensitive sensors for rapid electrical detection of Hg²⁺ with a detection limit as low as 30 pM.

Zhou *et al.* reported on a DNA-functionalized MoS₂ nanosheet/gold nanoparticle hybrid FET sensor for the ultrasensitive detection of Hg²⁺ in an aqueous environment (Figure 17a).²⁹⁷ A thin film was formed by filtration of MoS₂ nanosheets produced by liquid-phase exfoliation, followed by transfer onto the Au electrodes and thermal annealing. Then Au NPs were sputtered onto the MoS₂ film, and finally the DNA molecules were grafted onto Au NPs by immersing the device in a DNA solution. In the hybrid structure, the MoS₂ thin film acts as the conducting channel with the homogeneously dispersed Au NPs operating as anchoring sites for DNA probes chosen to specifically target Hg²⁺ ions. Upon addition of Hg²⁺ to the sensor, the formation of metal complexes between the thymidine bases present in the DNA molecules and mercury ions takes place, resulting in the changes of the MoS₂ conductance. The detection of metal ions was enabled by monitoring the change of the source-drain current in the FET device as a function of Hg²⁺ concentration. The detection limit of the sensor can reach values down to the concentration of 0.1 nM (Figure 17b). It was shown that the rate of the increase in the conductance or source-drain current is dependent on the Hg²⁺ concentration (Figure 17c). The sensor was also tested with a series addition of Hg²⁺ and the dynamic responses indicated that the sensor responded to Hg²⁺ within a few seconds, which is much faster than the conventional optical methods. The output characteristics of three types of sensors, i.e. MoS₂, MoS₂-Au NPs, MoS₂/DNA-Au NPs are shown in Figure 17d. Because of the work function difference between the Au NPs (3.6 eV in air for D = 8 nm) and the MoS₂ film (~5.23 eV after annealing), electron transfer occurs from the Au NPs to the MoS₂ film, leading to a decreased concentration of holes in the MoS₂ film and thus the decreased conductivity.

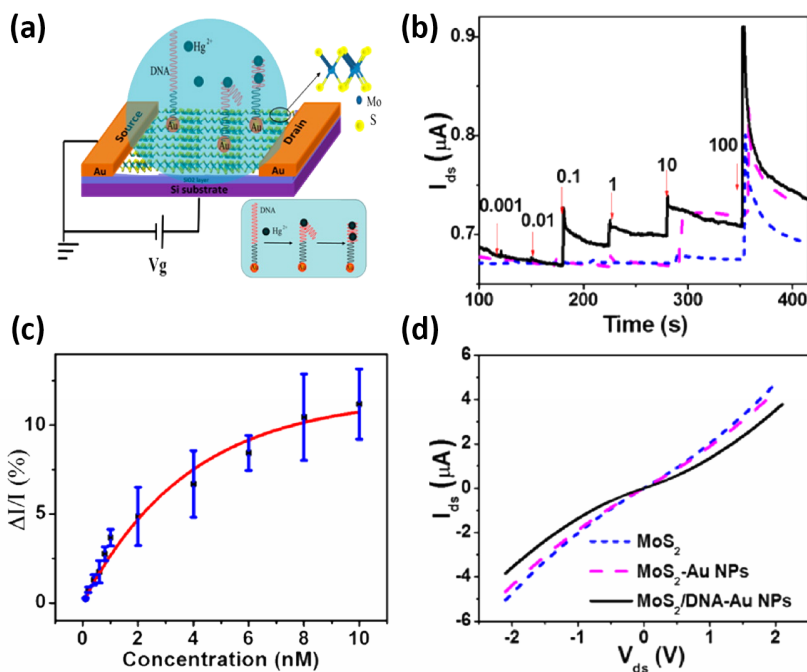


Figure 17. (a) FET sensor platform based on the hybrid structure. The formation of T-(Hg²⁺)-T chelates, through reaction between Hg²⁺ and the thymidine bases exposed on the DNA molecules grafted to the Au NPs, leading to the change in the MoS₂ electrical conductivity as a sensor signal. (b) Real-time detections of Hg²⁺ (nM) in water ($V_{ds} = 0.1$ V) with platforms of MoS₂/DNA-Au NPs (black, solid), MoS₂-Au NPs (purple, dash), MoS₂ (blue, short dash), respectively. (c) Sensitivity variation and exponential fitting of sensitivity as a function of Hg²⁺ concentration for the MoS₂/DNA-Au NPs hybrid sensor. (d) Evolution of the I_{ds} - V_{ds} characteristics during the MoS₂/DNA-Au NPs hybrid sensor fabrication process ($V_{ds} = -2.1$ V to 2.1 V, step = 0.1 V) at room temperature. Adapted from Ref. ²⁹⁷ with permission of American Chemical Society.

Despite the examples discussed in this Review, little knowledge has been gathered on the sorption capacity of other metal ions on 2D MoS₂. Therefore, it is of significant importance to extend in the future the application of such 2DM for the detection or removal of other heavy metals considering its markedly high metal capture capacity.

3.3.3 Other 2DMs based FET metal sensors

Li *et al* as the first presented ultra-sensitive suspended black phosphorus FET sensor for detection mercury ions operated in sub-threshold regime.²⁹⁸ Thin BP nanosheets were mechanically exfoliated by scotch tape based method and transferred onto low resistance Si substrate covered with a 300 nm thick layer of SiO₂. BP FET is a transducer, which converts the adsorption of ions into conductivity/current shift by gating effect. Free-standing BP FET sensors gated into sub-threshold regime demonstrated ultrahigh sensitivity which was almost two orders of magnitude larger than that of non-suspended BP in linear regime. The proposed BP sensors were able to realize rapid (3 s) label-free detection of Hg²⁺ down to 0.01 ppb, and superb selectivity was achieved by functionalizing BP surface with mercury ionophore. BP sensors tested with different thicknesses revealed that the response ($\Delta R/R_0$) to 10 ppb Hg²⁺ decreased approximately 30% when BP channel thickness varied from 8 nm to 20 nm

The same group demonstrated also air stable high-performance BP chemical sensors encapsulated with ionophore (Figure 18a, inset).³¹⁶ Without protection, BP samples once exposed to air start to degrade and oxygenated phosphorus (PO_x) is shortly formed, with degradation of its electric properties and device performances. Ionophore film effectively reduces negative factors from ambient environment; meanwhile, it displays selective permeability towards certain type of molecules. The ionophore-encapsulated BP devices were found being still in good shape after 1 week of ambient exposure, with source-drain current I_{DS} variation less than 10% (Figure 18a). The methodology of preparation thin BP nanosheets is the same as the one shown above (i.e. mechanical exfoliation with scotch tape). The BP sensors were sensitive to the detection of multiplex ions such as AsO²⁻, Hg²⁺, Cd²⁺, and Pb²⁺ with detection limit down to 10, 1, 3 and 1 ppb, respectively. Moreover, Pb²⁺ ions can be effectively detected over a wide concentration range, from 10 ppb to 100,000 ppb (Figure 18b). The detection limit and response rate of BP are both better than those reported for graphene based sensors.

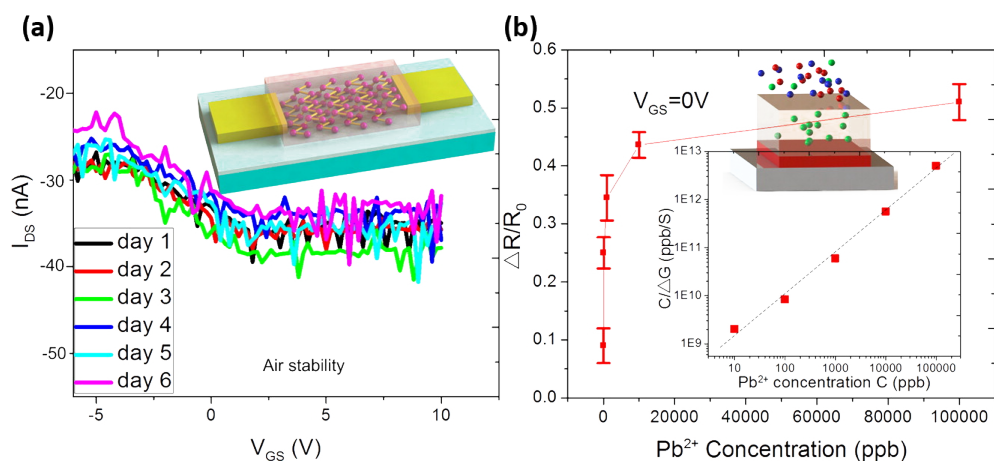


Figure 18. (a) Schematic view of a BP sensor with I_{DS} vs V_{GS} curves of BP with ionophore protection. I_{DS} variation is less than 10%, suggesting significantly improved air stability. (b) $\Delta R/R_0$ versus Pb^{2+} concentration (experimental results from 7 sensors) and schematic view of BP sensors with lead ionophore. Resistance of BP shows saturated at higher concentrations. Experimental results demonstrate good linear relation between $(C/\Delta G)$ and C , fitting well the Langmuir adsorption isotherm, as shown in the inset. Adapted from Ref. ³¹⁶ with permission of American Chemical Society.

Although only a few biosensors and water sensors with BP FETs were reported due to the inherent instability of BP in a moisture rich environment, BP sensors have attracted increasing interest across the sensing community due to the 2D structure, tuneable band gap, and high carrier mobility of BP. From a practical application standpoint, the thickness, uniform control, and stability of BP are limiting its application in ultrasensitive sensor technology. With further understanding of interactions between BP and oxygen/water and the improvement of sensor fabrication techniques, BP-based FET sensors can be foreseen to provide a robust sensing method for a wide variety of target analytes in the near future.

3.4 Electrochemical based metal sensors

Electrochemical sensing of heavy metal ions relies on the use of sensing electrodes that are employed for passing the current to the aqueous solution and generate electrical signal that corresponds to the electrochemical reaction within the solution due to presence of metal ions. Common experimental setups for electrochemical detection of heavy metal ions consist of an

electrolytic cell containing an electrolyte, i.e. solution of heavy metal ions, in contact with an electrode. The cell potential is measured at the interface of the electrode with the electrolyte solution. Noteworthy, because heavy metal ions have defined redox potential, the selectivity toward specific heavy metal ions can be achieved by using bare electrodes without the need of a molecular recognition probe. Numerous techniques are employed in electrochemical sensing, including potentiometry, voltammetry, impedimetry, amperometry and conductometry. In particular, the anodic stripping voltammetry (ASV) method is widely explored for detection of heavy metals. ASV analysis typically involves two steps, i.e. deposition of heavy metals onto the electrode surface, and stripping or dissolution of the deposited analyte from the electrode surface. Recent advancements in the field have revealed the potential of 2DMs in electroanalysis. Several electrochemical sensors based on 2DMs for bioanalysis and environmental analyses have been developed. In particular, Zhao *et al.* presented for the first time that Hg^{2+} can be selectively identified using a PPy-rGO nanocomposite-modified glassy carbon electrode (GCE).²⁹⁹ Such selectivity was achieved by using square wave anodic stripping voltammetry (SWASV), which determines a reduction of adsorbed Hg^{2+} to Hg^0 at a certain potential. The anodic stripping current was obtained in a potential range for the identification of Hg^{2+} . In addition, excellent sensitivity ($0.124 \mu\text{A nM}^{-1}$) and limit of detection (LOD) (15 nM) results were achieved. The measured stripping current toward Hg^{2+} at the PPy modified electrode was found to be 3–9 times higher than that towards other ions, indicating that rGO in the nanocomposite plays an important role in the highly selective detection.

Sahoo *et al.* reported a facile *in situ* approach for the fabrication rGO/bismuth(Bi) nanocomposite by employing modified Hummers method without the use of any surfactants. Bi nanoparticles were uniformly anchored onto the surfaces of individual graphene nanosheets, which prevent restacking of rGO, resulting in good dispersion in solvents. rGO/Bi nanocomposite was used as an electrode material for the stripping voltammetric determination of heavy metal ions in water. The detection limits of the proposed electrochemical sensor for Cd^{2+} , Pb^{2+} , Zn^{2+} and Cu^{2+} were found amounting to 2.8, 0.55, 17 and 26 ppb, respectively.³⁰⁰ Wei *et al.* developed a SnO_2/rGO -based electrochemical sensor, which could simultaneously and selectively analyse four heavy metal ions such as Cd(II), Pb(II), Cu(II) and Hg(II).³⁰¹ SWASV has been employed for the detection of heavy metal ions (Figure 19a). The SnO_2/rGO nanocomposite modified glass carbon electrode synthesized

by a simple wet chemical method showed enhanced sensing performance compared with single SnO₂ and single rGO (Figure 19b).

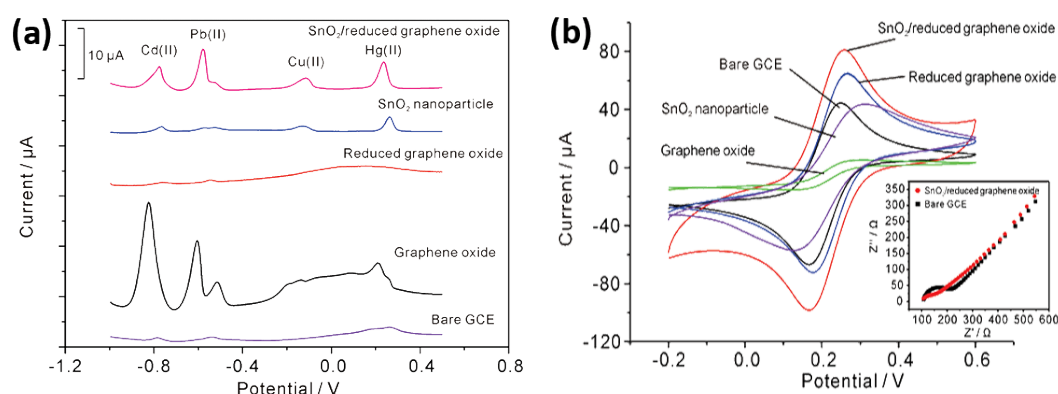


Figure 19. (a) Square Wave Anodic Stripping Voltammetry (SWASV) curves for 0.5 μM each of Cd(II), Pb(II), Cu(II), and Hg(II) on bare (violet line), graphene oxide (black line), reduced graphene oxide (red line), SnO₂ nanoparticle (blue line), and SnO₂/reduced graphene oxide nanocomposite (pink line) modified GCE in 0.1 M acetate buffer (pH 5.0). Deposition potential = -1.0 V; deposition time = 120 s; amplitude = 25 mV; increment potential = 4 mV; frequency = 15 Hz; vs. Ag/AgCl. (b) Cyclic voltammograms measured with bare, GO, rGO, SnO₂ nanoparticle, and SnO₂/reduced graphene oxide nanocomposite modified GCE in the solution of 5 mM Fe(CN)₆^{3-/4-} containing 0.1 M KCl. Inset: Nyquist diagram of electrochemical impedance spectra for bare and SnO₂/reduced graphene oxide nanocomposite modified GCE in the solution of 5 mM Fe(CN)₆^{3-/4-} containing 0.1 M KCl. Adapted from Ref. ³⁰¹ with permission of American Chemical Society.

Li *et al.* used Nafion and rGO for anodic stripping voltammetric analysis of cadmium with detection limit 0.005 ppb. ³⁰² Willemsse *et al.* reported on the determination of Cd²⁺, Pb²⁺, Zn²⁺ and Cu²⁺ by making use of the platform based on the Nafion-graphene nanocomposite film. ⁴⁰⁹ LODs of 0.07–0.08 ppb have been achieved for the individual ions, which are comparable to those determined with ICP-MS, and ascribed to a combination of enhanced electron conduction of rGO and the cation exchange capacity of Nafion. Similarly, Chaiyo *et al.* constructed a Nafion/ionic liquid/graphene electrochemical sensor for simultaneous determination of zinc, cadmium and lead using screen-printed carbon (Figure 20a). ³⁰³ The functionalized graphene-based nanocomposite modified electrode showed better detection performance for Zn(II), Cd(II) and Pb(II) compared to bare electrode by SWASV (Figure 20b). The detection limits of such sensors for Zn(II), Cd(II) and Pb(II) detection amounted to 0.09 ng mL⁻¹, 0.06 ng L⁻¹ and 0.08 ng L⁻¹, respectively.

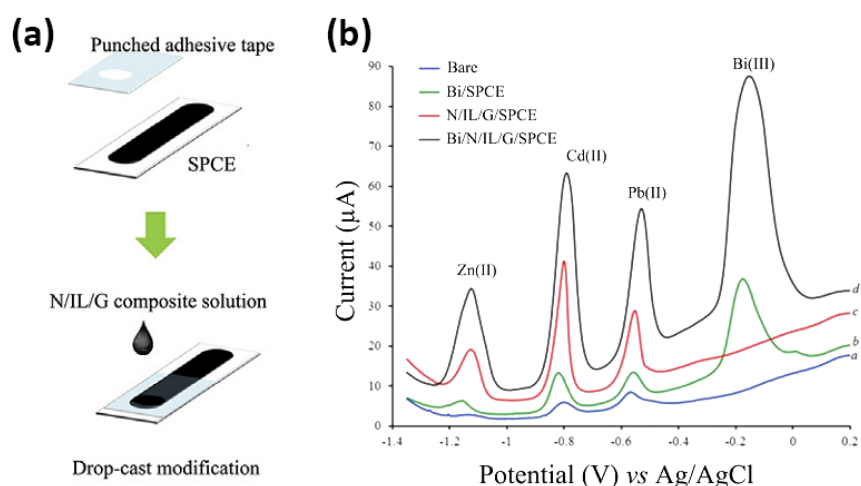


Figure 20. (a) Schematic drawing of the electrochemical sensor fabrication. (b) Square wave anodic stripping voltammetry (SWASV) curves of 50 ng mL^{-1} Zn(II), Cd(II) and Pb(II) in 0.1 M acetate buffer solution (pH 4.5). Adapted from Ref. ³⁰³ with permission of Elsevier B.V.

Gong *et al.* reported ultrasensitive Hg(II) electrochemical sensor by using monodispersed Au nanoparticles onto the graphene nanosheet matrix as the enhanced sensing platform.⁴¹⁰ The detection limit was found to be as low as 6 ppt. The interference from other heavy metal ions such as Cu^{2+} , Cr^{3+} , Co^{2+} , Fe^{3+} , Zn^{2+} and I^- ions associated with Hg^{2+} analysis could be effectively inhibited. Another example relies on the use of cysteine-functionalized GO (sGO) and carbonyldiimidazole as a cross-linker *via* the formation of amide and carbamate bonds. The sGO/polypyrrole (PPy) nanocomposite film was grown on the working electrode surface of a screen-printed electrode (SPE) *via* controlled one-step electrochemical deposition.³⁰⁴ The sGO/PPy-SPE was used to detect lead ions in water by differential pulse voltammetry (DPV). The DPV signals were linear in the ranges of 1.4–14 000 ppb of Pb^{2+} . The measurable detection limit of the sensor is 0.07 ppb, being 2 orders of magnitude below the threshold value for drinking water set by the World Health Organization. The average removal efficiency of Pb^{2+} deposited on the electrode amounted to 99.2%, with relative standard deviation (RSD) of 3.8%. The selective detection of the GO/PPy composites was investigated by mixing GO/PPy with various metal ions such as Pb^{2+} , Na^+ , Mg^{2+} , Cd^{2+} , Cu^{2+} , Hg^{2+} and Ag^+ . The results clearly show that the GO/PPy composites film can selective detection of

Pb²⁺ and impassive by the presence of other metal ions. In addition, the developed device can be used multiple times.

TMDs nanostructures were mainly utilized as sensing platforms for the development of electronic and fluorescent sensors, while their electrochemically sensing applications are still limited. Cui *et al.* synthesized monolayer MoS₂ supported Cu₇S₄-Au nanocomposite as a sensing platform to detect Hg(II) by using anodic stripping voltammetric technique.³⁰⁵ They found the synergistic effects of both Au domains and active edge sites of monolayer MoS₂ played a vital role in the high-performance for detection of Hg(II) with a limit of detection of 190 nM.

In this part of the Review article, we have provided an overview on the recent advances in the metal ion sensing. Different types of 2DMs employed for specific and selective recognition of heavy metals have been discussed in this chapter. By taking advantage of their unique properties 2DMs can be successfully explored to construct a wide range of optical and electrical sensing platforms for the detection of various heavy metal ions. Chemical modification of GO and TMDs has proven to provide promising solution for improving the sensing performance with high specificity, enhanced sensitivity and low detection limit. Chemical metal sensors based on 2DMs have demonstrated high sensitivity detection of a wide variety of heavy metal ions at low concentrations, due to the maximum sensor surface area per unit volume. Their favourable structural and compositional synergy allows them to be excellent electrode materials for fabricating various fluorescent and electrochemical sensing platforms, such as FET-based sensors. Moreover, electrochemical measurements have shown numerous advantages for trace heavy-metal detection, including rapid analysis, good selectivity, and sensitivity. Optical detection systems are other alternatives to the electrochemical detection methods. These represent attractive analytical tools whenever continuous monitoring and real-time information is desired.

4 (Bio)molecular sensors

4.1 (Bio)molecular sensing with graphene, graphene oxide and related composites

Ultrafast sensing of chemically and biologically active molecules at a low concentrations is critical in a wide range of research fields and in particular for applications, such as chemical analysis⁴¹¹ and healthcare,^{412, 413} monitoring the environment or diagnostic diseases.⁴¹⁴ Because of their particular physico-chemical properties, graphene-based materials have been used to fabricate various types of chemical sensors including electrochemical,^{72, 413, 415-444} FET,⁴⁴⁵⁻⁴⁴⁹ fluorescent⁴⁵⁰⁻⁴⁵⁷ and surface enhanced Raman spectroscopy-based sensors.^{411, 458-460} Graphene based materials have been used to develop various types of sensors to detect molecules such as glucose,^{72, 415-423, 425, 427, 445, 461-465}, DNA,^{72, 428, 430, 447, 449, 451, 452, 455, 457, 466-469} hydrazine,^{431, 470-473} dopamine,^{72, 413, 424, 432-439, 441, 454, 474, 475} ascorbic acid,^{72, 413, 434, 441, 476, 477} H₂O₂,^{72, 443, 478-480} or other aromatic molecules (Figure 21).^{411, 453, 456, 458, 459, 478} Graphene-based materials offer various advantages, when compared to other carbon-based nanomaterials like CNTs or fullerenes. In particular, graphene can be easily produced from bulk graphite *via* exfoliation.^{126, 481-483} Moreover, the particular strength of graphene, quality of its crystal structure, as well as its band structure and high conductivity allows the preparation of devices with extremely low noise levels and relatively low 1/f noise.⁴⁴⁸

4.1.1 Electrochemical (bio)molecular sensors

Electrochemical sensing is an effective and powerful technique used for qualitative and quantitative determination of bioactive and functional molecules.^{427, 465, 475} Numerous advantages including low limit of detection, high selectivity and detection in the presence of interferences made graphene-based materials appealing for fabrication of electrochemical sensors. Because of these reasons, electrochemical sensors belong to the most widespread graphene-based sensors in recent years. The most common approach relies on the use of electrodes based on graphene^{442, 477} or modified graphene-based nanostructures.^{441, 474, 478} Fast electron transfer and effective electrocatalytic activity of graphene allows the effective detection of target molecules while oxidation.⁴⁸⁴ Among its numerous physico-chemical properties, graphene has wide electrochemical potential window (ca. 2.5 V in phosphate

buffer)⁷² and is characterized by lower charge transfer resistance (comparing with glass electrode)⁷² and well characterized redox potentials against most common redox couples.⁴⁸⁵ Keeley *et al.* developed an electrochemical sensor based on graphene exfoliated in DMF for ascorbic acid (AA) detection.⁴⁷⁷ The electrode was prepared by drop-casting centrifuged graphene supernatant solution onto pyrolysed photoresist film. Cyclic voltammetry investigation revealed wide range of linearity and limit detection up to 0.12 mM. Another effective electrode was prepared by mixing platinum nanoparticle and graphene nanosheet.⁴⁷⁸ Such hybrid system displayed the capacity to detect traces of trinitrotoluene (TNT), a typical explosive compound, with detection limits as low as 0.3 ppm and satisfactory reproducibility. Luo *et al.* presented electrochemical glucose sensor based on Cu decorated graphene sheets.⁴²¹ As prepared electrode exhibited a low detection limit of 0.5 μM and very fast response ($<2\text{s}$) and its response linearity up to 4.5 mM. Additionally, a synergistic effect of copper and graphene toward glucose oxidation was observed.

As aforementioned, the majority of graphene-based sensors rely on the chemical modification of GO, which exhibits numerous oxygen-rich functional groups that can interact *via* dipole-dipole or strong electrostatic interactions with (charged) molecules, or it can be covalently functionalized with functional molecules, enhancing the occurrence of adsorption events. In particular, covalent grafting of organic molecules on GO offers the flexibility for various functionalizations to enhance the sensor performance. Moreover, the combination of its abundant structural defects and chemical groups facilitates the charge transfer and thus ensure high electrochemical activity. Noteworthy, the chemical and electrical properties of rGO are highly tuneable and can be engineered through control of the reduction process.

Zhou *et al.* exploited a chemically reduced GO modified glassy carbon electrode (CR-GO/GC) for electrochemical sensing of free purine and pyrimidine acids and several other biologically active molecules.⁷² It was demonstrated that CR-GO/GC electrode provides greater electrocatalytic activity to guanine (G), adenine (A), thymine (T) and cytosine (C) oxidation than commonly used glassy carbon (GC) and graphite/glassy carbon graphite/GC electrodes (Figure 21). Moreover, the analytical performance of glucose, ethanol, dopamine and other biological molecules detection was enhanced; e. g. compared with graphite/GC and GC electrodes the CR-GO/GC electrode. In particular, while amperometric measurement of reduced nicotinamide adenine dinucleotide (NADH) exhibited faster response time, wider linear range (40-800 μM), lower detection limit (10.00 μM) and higher sensitivity (2.68 $\mu\text{A mM}^{-1} \text{cm}^{-2}$). Gao and co-workers carried out the chemical reduction of GO obtaining

hydroxyapatite/rGO composite used as effective electrochemical sensor of hydrazine.⁴⁷² The composite dispersed in 1% acetic acid was applied on GCE and dried resulting in high-performance electrode with outstanding performance of electrocatalytic oxidation of N₂H₄. The as-prepared electrode exhibited synergic effect of components during detection of hydrazine, being greater than the electrodes fabricated with rGO and hydroxyapatite separately. Wang and co-workers devised a facile approach relying on the efficient preparation of nitrogen-doped graphene *via* nitrogen plasma treatment of chemically synthesized graphene, which exhibits excellent electrocatalytic activity toward hydrogen peroxide reduction.⁴¹⁵ Detection of H₂O₂ as one of the products of catalytic glucose oxidation can be applied for selective and sensitive glucose determination during enzymatic processes. Noteworthy, N-doped graphene electrode exhibited high concentration-dependent response and was able to detect up to 0.01 mM of glucose in the presence of interferences. N-doped electrodes can also find their use in detection of important biomolecules such as ascorbic acid (AA), dopamine (DA) and uric acid (UA).⁴¹³ Efficient oxidation and large peak separation provides simultaneous determination of those three biomolecules with detection limits up to $2.2 \cdot 10^{-6}$, $2.5 \cdot 10^{-7}$ and $4.5 \cdot 10^{-8}$ M for AA, DA and UA, respectively. Furthermore, differential pulse voltammetry (DPV) experiment revealed that electrochemical response of target biomolecules increases linearly with increase of molecules concentrations.

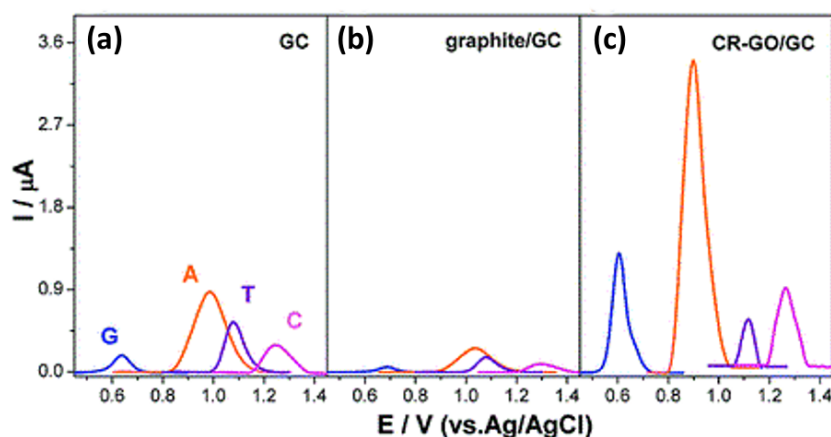


Figure 21. Differential pulse voltammograms (DPVs) recorded by using a) GC electrode, b) graphite/GC electrode, and c) CR-GO/GC electrode, for guanine (blue), adenine (orange), thymine (violet), and cytosine (magenta), respectively. Adapted from Ref. ⁷² with permission of American Chemical Society.

4.1.2 Field-Effect Transistors for (bio)molecular sensors

Due to their high selectivity, fast response and excellent limit of detection, graphene-based FET-based sensors became very popular for label-free ultrasensitive biomolecule sensing. By and large, graphene-based materials are used as conducting component in the channel between drain and source electrodes, and subsequently gate potential is applied.⁴⁴⁶ Two different approaches are being pursued for construction of FET sensors, i.e., back gating (usually Si/SiO₂ thin layer)^{445, 446} or top gating (electric double layer in electrolyte)^{449, 465} also known as solution gated graphene transistors (SGGTs). Both methods have been schematically presented in Figure 22. In general, charged molecules can induce an effective gating field, which influences the conductance of the channel by balancing the charge transfer and gating effect. The change in current as measured allows the analyte determination.

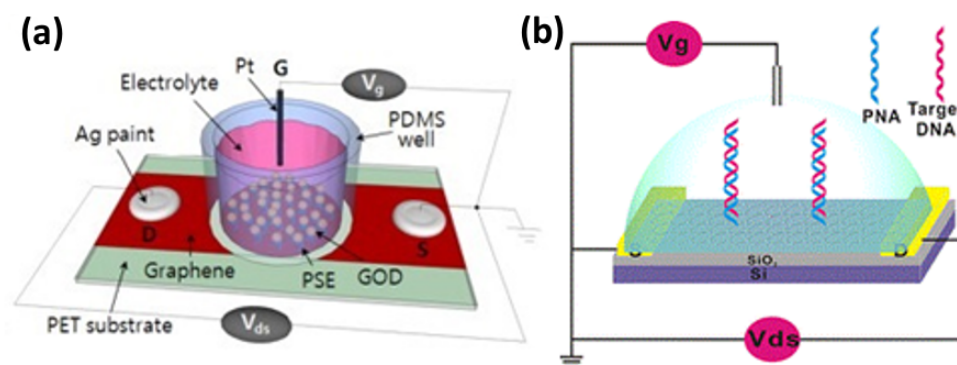


Figure 22. Schematic representation: (a) Graphene top gate FET sensor. Adopted from Ref. ⁴⁴⁶ with permission of the American Chemical Society. (b) Graphene oxide back-gate FET sensor. Adapted from Ref. ⁴⁴⁹ with permission of Elsevier.

Huang and co-workers presented CVD-grown graphene FET sensor functionalized with specific redox mediators for glucose and glutamate detection.⁴⁴⁵ Glucose sensing is usually based on an enzymatic reaction catalysed by glucose oxidase. Since the products of oxidation process are H₂O₂ and gluconic acid, direct measurement of H₂O₂ is useful for glucose detection. The detection limits were found being at 0.1 mM and 5 μM for glucose and glutamate, respectively. Another CVD-grown glucose sensor was assembled by Kwak *et al.*⁴⁴⁶ Solution-gated field-effect transistor (SGFET) was constructed utilizing graphene channel modulated by the gate potential applied from the top gate electrode and transmitted through the solution. Such SGFET sensor was exploited to detect glucose in the range of 3.3-

10.9 mM. Moreover, it provided high resolution and continuous real-time monitoring. To reach the largest sensing response, graphene transistors are operated at the point of maximum transconductance, which leads to large noise that influences the device sensitivity. To avoid such phenomenon Fu and co-workers exploited the sensing properties of single layered graphene near its neutrality point.⁴⁴⁷ This approach led to a significant decrease of signal-to-noise ratio, thereby making it possible to observe positive signals coming from single stranded DNA (ssDNA) at level of picomolar concentrations (pM). Moreover, to target specific hybridization corresponding with HIV-virus related ssDNA, functionalization of GFET surface with pyrene-linked peptide nucleic acid (pPNA) was performed. Noteworthy, the 1/f noise in graphene based FET sensors was found to vary with the number of graphene layers, therefore subsequent optimization of this class of sensors is anticipated.⁴⁴⁸ Another approach in which hydrazine-reduced GO was used to fabricate FET sensor on SiO₂/Si substrate was proposed by Cai *et al.*⁴⁴⁹ DNA sensor was prepared by drop-casting rGO suspension onto sensing channel as the conducting material. 1-Pyrenebutanoic acid succinimidyl ester (PASE) used as molecular linker was fixed on the graphene surface *via* π - π stacking interactions and peptide nucleic acid (PNA) molecules were covalently anchored. Subsequently, a complementary DNA was applied onto device and specific hybridization caused the shift of I_{ds} - V_g curves and allowed the measurement. The detection limit was estimated being as low as 100 fM.

4.1.3 (Bio)molecular Fluorescence Sensors

The use of fluorescence as readout in the detection of chemical species and biomolecules holds potential for low-cost, effective and highly sensitive device applications. Graphene/graphene oxide are known to be excellent quenching materials therefore graphene-based materials became suitable for sensing *via* influence on fluorescence properties.⁴⁶⁹ One of the approaches bases on FRET and is useful for quantitative determination of target molecules. Typically, this kind of set up consists of a donor, an acceptor and a bridge. The acceptor such as GO can effectively quench fluorescence of the donor. The addition of specific molecules or surface modification may cause change in the fluorescence intensity and can be utilized as a sensor assuming the phenomenon is correlated with concentration (Figure 23a-b).^{450, 452}

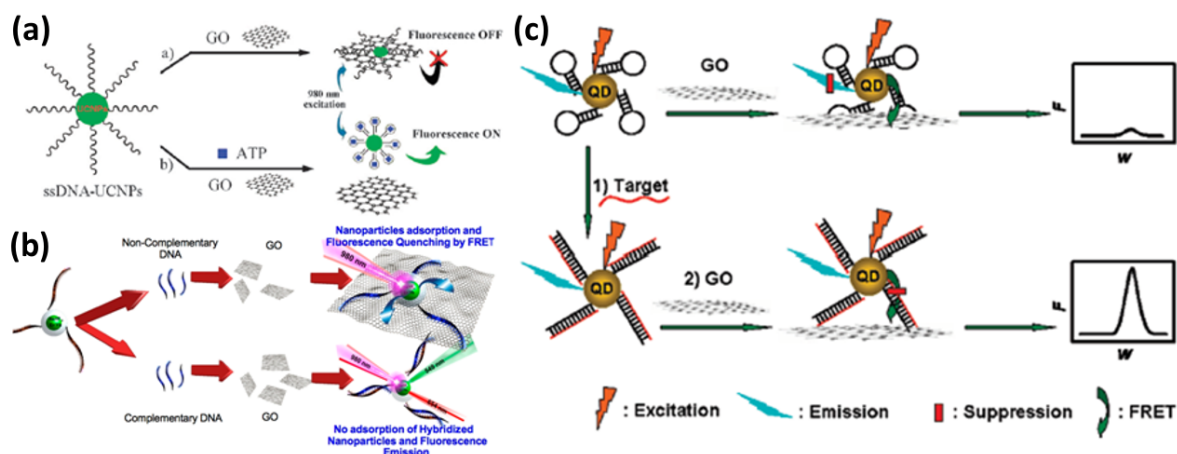


Figure 23. Schematic representation of different FRET mechanisms used in biomolecules sensing. (a) Schematic representation of the upconversion fluorescence resonance energy transfer used for ATP sensing. (Figure adapted from Ref. ⁴⁵² with permission of the Royal Society of Chemistry.) (b) Depending on the presence of complementary DNA emission is observed or completely quenched, detection of specific DNA hybridization. Adapted from Ref.⁴⁵⁰ with permission of American Chemical Society. (c) DNA sensing. Molecular beacon decorated QDs quenched by GO platform, in the presence of complementary DNA emission is significantly enhanced. (Figure adapted from Ref. ⁴⁵¹ with permission of American Chemical Society.)

Dong and co-workers reported on the FRET from quantum dots (QDs) to GO.⁴⁵¹ In particular, the authors demonstrated that QDs substituted with molecular beacon (MB) as a recognition unit towards targeted molecule can strongly interact with GO surface, to allow the design of novel sensitive and selective platforms for fluorescence-quenching detection of DNA. In general GO's quenching properties were used to decrease fluorescence intensity in the presence of QDs – an effective donor supplied with MB that provides efficient energy transfer to GO (Figure 23c). Fluorescent graphene quantum dots (GQDs) were also applied for detection of aromatic nitro compounds (e.g. TNT) working as a quenchers in a wide range sustaining linear correlation.⁴⁵³ Another DNA sensor using FRET between upconversion nanoparticles and GO was reported by Alonso-Cristobal *et al.*⁴⁵⁰ Correlation between concentration of ssDNA (Figure 24) and luminescence intensity was presented. Effective detection limit was experimentally shown to be in the picomolar range.

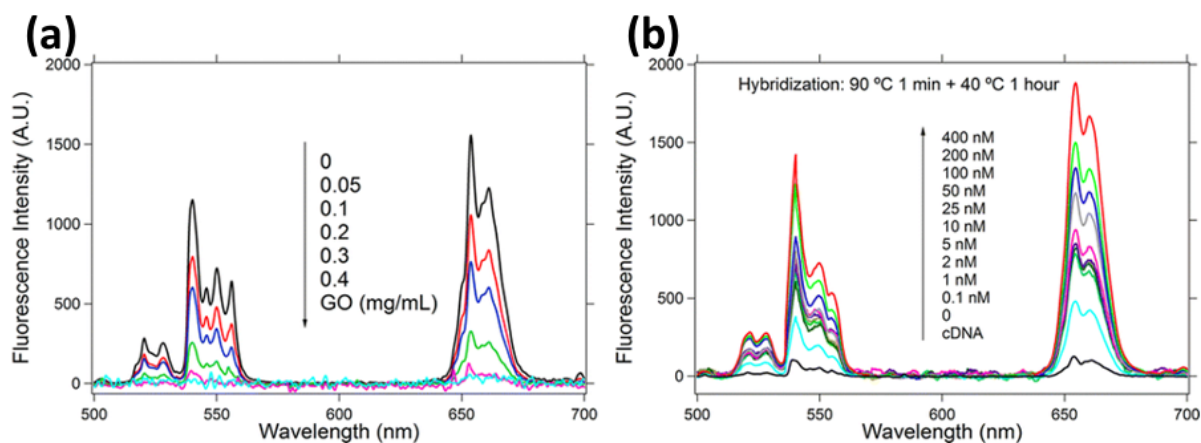


Figure 24. Up-conversion fluorescence spectra of the UCNPs@SiO₂-ssDNA nanoparticles (0.4 mg mL⁻¹) a) in the presence of different concentrations of GO, and b) in the presence of different concentrations of complementary ssDNA and 0.3 mg mL⁻¹ of GO. Adapted from Ref.⁴⁵⁰ with permission of American Chemical Society.

GO-based platform for DNA and proteins sensing reported by Lu and co-workers includes ssDNA supplied with fluorescent organic dye that strongly adsorbs on GO's surface simultaneously entailing high quenching efficiency.⁴⁵⁵ The addition of complementary ssDNA sequence causes efficient fluorescence increase and may be used to determine desired nucleotide sequences. Experiments also proved the dependence on different proteins concentrations. The results revealed that the presence of human thrombin might be determined without the interference of other proteins.

4.1.4 (Bio)molecular Sensing *via* Surface Enhanced Raman Scattering

Surface Enhanced Raman Scattering represents a most powerful analytical technique for sensitive chemical sensing, which became an attractive solution for detection of aromatic compounds. Typically, weak Raman signals are intensified through enhancement of the local electromagnetic field, occurring with the charge transfer between adsorbed molecules and metal surface. Consequently, to observe SERS signal, molecules must be adsorbed or be placed close to a metal surface. The most common approach bases on modification of graphene-based material surface with metallic nanoparticles (NPs).^{411, 459} Optimum size,

shape, charge and surface chemistry of NPs are being taken into account to project non-toxic and effective sensors.¹⁰⁴

Lu *et al.* proposed the use of Ag and Au NP-decorated single-layer rGO film anchored on Si surface to detect popular organic dyes rhodamine 6G (R6G), methyl violet (MV), rhodamine B (RB) and methylene blue (MB) at nanomolar concentrations.⁴¹¹ Such molecules easily adsorb on rGO surface while Ag and Au nanoparticles provide dramatic enhancement in the Raman intensities coming from the sample. Moreover, a correlation between NPs size and observed signals have been monitored, and it could be controlled through the reaction time. Another modification with Ag NPs for ultrasensitive detection of aromatic molecules was demonstrated by Liu and co-workers.⁴⁵⁸ Additional enhancement of the Raman signals was ensured by GO treated with 3-mercaptopropyltrimethoxysilane (MPTMS), further functionalized with Ag-NPs. An effective detection of differently charged molecules was successfully characterized and presented. In particular, experiments with PPh₃ used as target molecule exhibited limit of detection as low as 10⁻⁹ M.

Graphene grown on copper foil and Ag NPs anchored on Si substrate can also be used to prepare SERS active substrate for detection of TNT by alkaline hydrolysis.⁴⁵⁹ High pH values provide efficient hydrolysis of TNT and enhance intensity in Raman spectra signals. Extremely high limit of detection (up to 6.6·10⁻¹⁰ M) and excellent anti-interference ability confirms perspective potential of SERS sensors. Shanta and Cheng prepared hydrazine reduced GO combined with silver nanoprisms and utilized such hybrid system for trace detection of tetrachlorobiphenyls (PCBs).⁴⁶⁰ As prepared chip successfully allowed for multiplex measurement of several environmentally important aromatic compounds. Proposed SERS sensor provided low limit of detection (up to 100 nM) and simultaneous analysis of multiple isomers.

Values such as limit of detection and linearity range for detection of vary biomolecules have been summarized in Table 9.

Table 9. Comparison of sensing properties of different graphene-based sensors.

<i>Material</i>	<i>Type of sensor</i>	<i>Analyte</i>	<i>Limit of detection</i>	<i>Linearity Range</i>	<i>Ref.</i>
<i>DMF-exfoliated</i>	electrochemical	AA	0.12 mM	0.4-6.0 mM	⁴⁷⁷

<i>graphene</i>				
<i>Pt-NPs-graphene</i>		H ₂ O ₂	80 nM	1-500 μM
		TNT	0.3 ppm	0.5-40 ppm
<i>Cu-NPs-graphene</i>		glucose	0.5 μM	0-4.5 mM
<i>rGO</i>		H ₂ O ₂	0.05 μM	0.05-1500 μM
		NADH	10 μM	40-800 μM
<i>Hydroxyapatite-rGO</i>		NH ₂ NH ₂	0.43 μM	2.5 μM - 1.16 mM
<i>N-doped graphene</i>		glucose	0.01 mM	0.1-1.1 mM
<i>N-doped graphene</i>		AA	2.2 μM	5.0-1300 μM
		DA	0.25 μM	0.5-170 μM
		UA	4.5×10 ⁻⁸ M	0.1-20 μM
<i>CVD-grown graphene</i>	FET	glucose	0.1 mM	-
		glutamate	5 μM	5 - 1200 μM
<i>CVD-grown graphene</i>		glucose	-	3.3-10.9 mM
<i>CVD-grown graphene</i>		DNA	2 pm	-
<i>rGO</i>		DNA	100 fM	10 fM-1 nM
<i>QDs-GO</i>	fluorescent	DNA	12 nM	50-1500 nM
<i>Graphene QDs</i>		TNT	2.2 μM	4.95×10 ⁻⁴ - 1.82×10 ⁻¹ g/L
<i>NPs-GO</i>		DNA	5 pm	-
<i>Au/Ag-NPs-rGO</i>	SERS	R6G		
		MV	nm	-
		RB		
		MB		
<i>Ag NPs-GO</i>		PPh ₃	1 nM	5×10 ⁻⁶ -1×10 ⁻⁹ M
<i>Ag NPs-graphene</i>		TNT	0.66 nM	1-100 nM
<i>rGO-NP</i>		R6G	100 pM	0-1000 nM
		PCB	100 nM	100 μM - 100 nM

List of abbreviations. DMF : dimethylformamide, AA : ascorbic acid, NPs : nanoparticles, TNT : 2,4,6-trinitrotoluene, NADH : β-nicotinamide adenine dinucleotide, NH₂NH₂ : hydrazine, DA : dopamine, UA : uric acid, QDs : quantum dots, SERS : Surface Enhanced Raman Scattering, R6G : rhodamine 6G, MV : methyl violet, RB : rhodamine B, MB : methylene blue, PPh₃ : triphenylphosphine, PCB : 2,2',3,3'-tetrachlorobiphenyl.

4.2 (Bio)molecular sensing with Transition Metal Dichalcogenides

TMDs were recently used to develop a wide range of chemical sensor including electrochemical,^{132, 134, 486-500} fluorescent^{99, 133, 501-503} and field-effect transistor sensors.⁵⁰⁴⁻⁵⁰⁶ MoS₂ and its analogues like tungsten disulphide (WS₂), tantalum disulphide (TaS₂), tungsten diselenide (WSe₂) as well as ternary chalcogenides are widely investigated with promising results in chemical sensing applications.^{132, 500-502, 507}

Among various TMDs, MoS₂ is the most studied material. The interaction of 2D lattice of MoS₂ with charged molecules efficiently influences its electrochemical properties, which can be measured directly with voltammetric measurements.^{490, 492, 493} Furthermore, MoS₂ nanosheets used for electrode preparation exhibit significant electrocatalytic activity toward a range of important analytes, including glucose,^{493, 497} H₂O₂,⁴⁸⁸ dopamine,^{494, 495, 508} ascorbic acid,⁵⁰⁸ tryptophan⁵⁰⁹ and DNA.^{489, 503, 506} Another advantage of MoS₂ is its feasibility to form composites which enlarges a wide range of properties and application potential.^{509, 510} Remarkably, the presence of noble metal nanostructures effectively enhances electrochemical detection of target molecules. Upon mixing MoS₂ with of NPs (e.g. Pt, Ag, Au) improves electron transfer and allows achieving lower limits of analyte detection.^{495, 508, 509} MoS₂ is also an excellent quenching material, thus can be applied for highly sensitive detection of biomolecules *via* fluorescence phenomenon.^{99, 511} Moreover MoS₂ occurs high affinity toward ssDNA, which is supported by van der Waals interactions, while double stranded DNA (dsDNA) formed in the presence of complementary sequence, hardly interacts with its two-dimensional surface. Different mechanisms of molecular detection utilizing TMD's based materials are presented on Figure 25.

Wu and co-workers presented controllable lithiation process of MoS₂ and subsequent liquid-phase exfoliation of the intercalated structure in water or ethanol. The electrochemical study of as-prepared MoS₂ nanosheets revealed a reduction peak in the CV in NaCl aqueous solution. The reduced MoS₂ displayed good conductivity and fast electron transfer rate in the most common redox systems applicable in electrochemical sensing of glucose and dopamine (DA) in the presence of ascorbic acid (AA) and uric acid (UA).⁴⁸⁶ GCE was covered with MoS₂ using 3-aminopropyltriethoxysilane (APTES) as linking molecule. The electrode prepared with such hybrid material yielding voltograms displaying good peak separation, thus allowing the detection of dopamine maintaining linearity in the range of 1-50 μM in the

presence of interferences (AA) and (UA). Additional functionalization of electrode with glucose oxidase enables effective glucose detection.

Another approach for molecular detection was introduced by Huang *et al.*⁴⁹⁷ MoS₂ nanosheet obtained *via* liquid-phase exfoliation in ethanol/water mixed with Ni NPs and drop-cast onto GCE and naftion copolymer. The as-prepared sensors exhibited amperometric response toward glucose with detection limit as low as 0.31 μ M. Such GCE/Ni-MoS₂/Naftion electrode allowed nonenzymatic detection of glucose with good reproducibility and long-term storage stability. The simultaneous detection of DA, AA and UA investigated by Sun *et al.* by using MoS₂-Au NPs decorated electrode to achieve ultrasensitive determination of aforementioned biologically important molecules.⁵⁰⁸ The joint effect of a wide linearity range and effective electrical response yielded sensing ability with detection limits as low as 100 μ M, 0.05 μ M and 10 μ M for AA, DA and UA, respectively. Zhu and co-workers have shown the fluorescence quenching ability of MoS₂.⁹⁹ Nanosheets obtained through exfoliation of bulk MoS₂ using electrochemical lithium-intercalation method were used as nanoprobe for adenosine. MoS₂ supplied with probe molecule occurs high fluorescence quenching efficiency (up to 98% comparing with probe molecule intensity). Interaction with adenosine observed as intensity increase can be used as highly selective sensor able to detect up to 5 μ M of adenosine. Another MoS₂ fluorescent sensor for DNA detection was proposed by Huang *et al.*⁵⁰³ The nanosheets were prepared by using solution exfoliation method. The authors presented selective target DNA recognition *via* hybridization chain reactions (HCRs) with stunning detection limit of 15 pM. Another example of using adsorption of dye-labelled ssDNA on MoS₂ surface was presented by Huang *et al.* Chemically exfoliated MoS₂ was utilized as biosensing platform for complementary DNA and allowed detection up to 500 pM.⁵¹² Shan *et al.* exploited MoS₂-based FET for glucose detection.⁵¹³ The device was fabricated on Si/SiO₂ substrate with Au/Ni drain electrodes. The sensor exhibited very fast response time (<1 s) and linearity in the range of 300 nM - 30 mM, which corresponds to average glucose concentration present in the blood consisting potential candidate for real-time detection of glucose concentration in human blood.

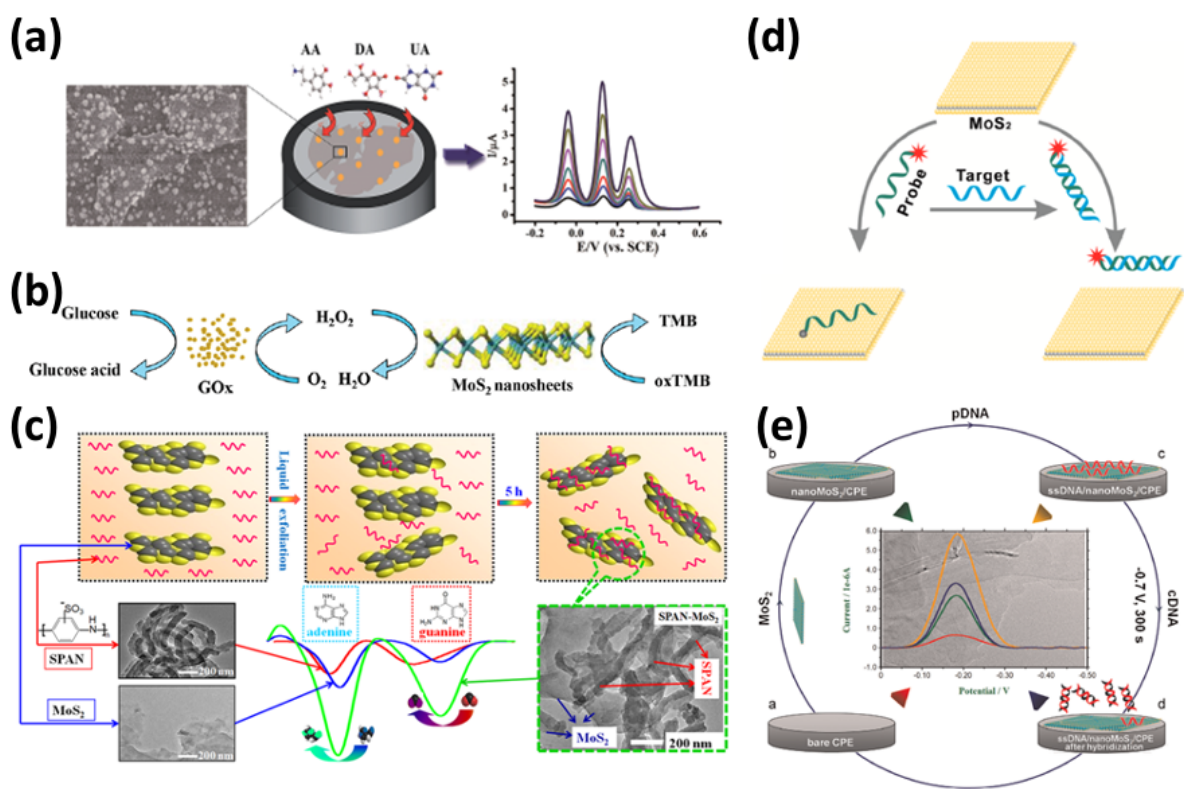


Figure 25. Representation of different mechanisms used for chemical detection utilizing MoS₂-based materials. (a) Electrochemical sensing of ascorbic acid (AA), dopamine (DA) and uric acid (UA) with gold nanoparticle-MoS₂ nanosheets. Reproduced with permission.⁵⁰⁸ Copyright 2014, Royal Society of Chemistry. (b) Sketch of catalytic reactions while colorimetric detection of glucose. Adapted from Ref.⁴⁹⁹ with permission of the Royal Society of Chemistry. (c) SPAN-MoS₂ for electrochemical detection of adenine and guanine. (Figure adapted from Ref.¹³⁴ with permission of the American Chemical Society). (d) Fluorescent detection of complementary ssDNA. (Figure adapted from Ref.⁹⁹ with permission of the American Chemical Society). (e) Label-free electrochemical sensing of DNA. Adapted from Ref.⁴⁸⁹ with permission of Elsevier.

Inspired by the unique device properties offered by MoS₂, other TMDs (including WS₂, TaS₂, Ta₂NiS₅) have been the subject of further studies for sensing applications. The group of layered TMCs, including the chalcogenides with well-defined crystal structures have been widely explored.^{502, 514} Since the MoS₂ is known for its quenching properties, the entire group of TMDs is being considered as effective platform for fluorescent sensing of biomolecules.^{501, 507}

Zhang *et al.* have used single-layer TMDs nanosheets as fluorescent sensors for DNA detection.⁵⁰¹ To investigate the quenching ability of TiS_2 and TaS_2 , 6-carboxyfluorescein-labeled ssDNA probe and target DNA molecule were used. In both cases the dichalcogenides fluorescence intensity was effectively quenched. TiS_2 as well as TaS_2 exhibited fast detection of complementary DNA with low limit of detection of 0.2 nM and 0.05 nM, respectively. Noteworthy, TaS_2 -based nanosensor was found to be even more sensitive than MoS_2 with detection limit of 0.1 nM. Another fluorescent sensing platform based on WS_2 for T4-polynucleotide kinase and its inhibitors was introduced by Ge and co-workers.¹³² The dsDNA does not exhibit strong adsorption on WS_2 sheets, yet when phosphorylated and degraded the ssDNA binds to nanosheets as evidenced by the occurrence of fluorescence quenching. Linearity, low limit of detection and low signal-to-background ratio make this material promising for application in monitoring of biological processes. Huang *et al.* developed WS_2 ink by sonication-assisted liquid exfoliation in water supported by Au nanoparticles for detection of 17β -estradiol.⁴⁹⁶ Electrode prepared by covering GCE by WS_2 , then Au NPs were anchored and ssDNA aptamer was applied for selective detection of target molecule. As prepared electrode allowed to determine 17β -estradiol at concentration as low as 2 pM. Moreover they proved that only target molecule may cause such obvious current change at concentration in range of 10^{-10} M. Tan and co-workers prepared single-layer ternary chalcogenide for fluorescence sensing of DNA.⁵⁰² Microsized crystalline flakes of Ta_2NiS_5 were obtained *via* chemical vapour transport technique from their elementary powders. Then high yield exfoliation was performed to produce single-layered compound. Probe DNA was supplied with Texas red fluorescent reporter functionalized with the probe DNA was employed since it is known to be efficiently quenched when interaction with TMDCs. The fluorescence could be recovered when the target ssDNA was added to form dsDNA. The fabricated sensor exhibited high sensitivity (up to 50 pM) and affordable detection time of 10 min. Ta_2NiS_5 was also used as “Capture-Release” nanosensor for detection of biological species (Figure 26).

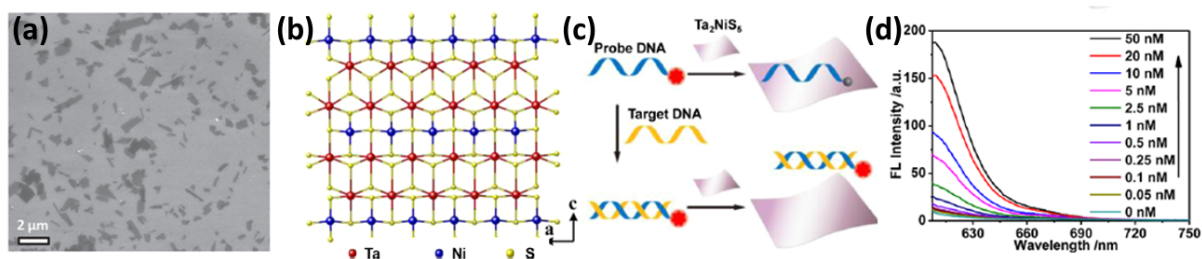


Figure 26. a) Scanning Electron Microscopy (SEM) image of Ta_2NiS_5 nanosheets. b) Crystal structure of Ta_2NiS_5 along the b-axis. c) Schematic representation of quenching ability of Ta_2NiS_5 applied for ssDNA detection. d) Fluorescence detection of different concentrations of target molecule in the presence of probe ssDNA. Adapted from Ref. ⁵⁰² with permission of American Chemical Society.

4.3 (Bio)molecular sensing with TMDs-graphene hybrid materials

Both TMDs and graphene exhibit a range of physico-chemical properties suitable for various electronic and sensing applications. In particular, further modifications of MoS_2 and generation of hybrid materials might be used to generate composites with desired properties. The synergistic effect on electrochemical performance of carbon-based materials with metal or metal oxides have been previously reported⁵¹⁵ thus suggesting that the combination of MoS_2 with graphene-like structures as excellent candidate for a new class of electrochemical sensors. Electrodes made of TMC-graphene composites exhibit significant enhancement of electron transfer conductivity.⁵¹⁶⁻⁵¹⁸

Huang *et al.* reported a MoS_2 -graphene composite prepared *via* modified L-cysteine-assisted solution-phase method casted on GCE to assemble an electrode.⁵¹¹ The composite displayed a wide electrochemical potential window (ca. 2.9 V in phosphate buffer), which is comparable to graphene-based electrodes. The as-prepared electrode was utilized for voltammetric determination of acetaminophen exhibiting linear response in the range of 1-100 μM . Moreover, the hybrid material showed good interference resistance: in the presence of other electrochemically active molecules the value of oxidation peak current was sustained. Another electrochemical sensor composed of MoS_2 -graphene nanohybrid film was developed by Zhao and co-workers.⁵¹⁹ The GCE was embedded in MoS_2 -graphene nanohybrid resulting in significantly enhanced sensitivity towards honokiol - known in view of its antitumor and antibacterial activity used in western medicine. The electrode revealed high selectivity and

stability with a spectacular limit of detection corresponding to 6.2×10^{-10} M. Pramoda *et al.* performed electrocatalytic activity of MoS₂-rGO/GCE against dopamine and proved its superior nature over bare rGO-modified electrode.⁵²⁰ Effective detection of DA in the presence of interferences was established at the level of 0.55 μ M.

The most important parameters, including type of sensor, limit of detection and linearity range for TMDs-based sensors are being presented in Table 10.

Table 10. Comparison of sensing properties of different transition metal chalcogenides based sensors.

<i>Material</i>	<i>Type of sensor</i>	<i>Analyte</i>	<i>Limit of detection</i>	<i>Linearity Range</i>	<i>Ref.</i>
<i>MoS₂</i>	electrochemical	H ₂ O ₂	2.5 nM	0.1-100 μ M	488
		DNA	0.019 fM	0.1 - 1×10^5 nM	489
	fluorescent	DNA	15 pM	0-200 pM	503
		DNA	500 pM	0-50 nM	512
	FET	glucose	300 nM	300 nM - 30 mM	513
<i>MoS₂-Pt NPs</i>	electrochemical	DA	0.17 μ M	0.5-150 μ M	495
		AA	0.98 μ M	5-1000 μ M	521
		H ₂ O ₂	1.0 μ M	4.0-48500 μ M	497
<i>MoS₂-NiNPs</i>		glucose	0.31 μ M	0-4 mM	509
<i>MoS₂-Ag/chitosan</i>		tryptophan	0.05 μ M	0.5-120 μ M	508
		AA	100 μ M	1-70 mM	494
<i>MoS₂-Au NPs</i>		DA	0.05 μ M	0.05- 4×10^4 μ M	494
		UA	10 μ M	10- 10^3 μ M	493
		DA	80 nM	0.1-200 μ M	493
		glucose	2.8 μ M	10-300 μ M	134
<i>MoS₂-PANI</i>		adenine	3 nM	0.05-100 μ M	498
		guanine	5 nM	0.05-100 μ M	500
		chloramphenicol	65 nM	0.1-1000 μ M	496
<i>WS₂</i>		glucose	2.9 μ M	5-300 μ M	522
<i>WS₂-Au NPs</i>		17 β -estradiol	1 pM	0.01-10 nM	501
<i>WS₂-acetylene black</i>		DNA	0.12 fM	0.001-100 pM	502
<i>TaS₂</i>	fluorescent	DNA	0.05 nM	0-5 nM	511
<i>TaNiS₅</i>		DNA	50 pM	0-5 nM	519
<i>MoS₂-graphene</i>	electrochemical	acetaminophen	20 nM	0.1-100 μ M	523
		honokiol	0.62 nM	1.0 nM -2.5 μ M	511
		H ₂ O ₂	1.25 μ M	6.25-225 μ M	523

<i>MoS₂-rGO</i>	DA	0.94 μ M	1.5-100 μ M	520
<i>HPR-MoS₂-graphene</i>	H ₂ O ₂	0.049 μ M	0.2 μ M -1.1 mM	510

List of abbreviations. NPs : nanoparticles, DA : dopamine, AA : ascorbic acid, UA : uric acid, PANI : polyaniline, TaS₂ : tantalum(IV) sulfide, TaNiS₅ : tantalum nickel sulfide, HPR : horseradish peroxidase.

4.4 (Bio)molecular sensing with other 2DMs

Hexagonal BN is known to be the most stable allotrope among BN crystalline forms. A few electron transfer reactions toward biomolecules oxidation for commonly used electrodes immobilized with boron nitride were reported and the results are comparable to other 2DMs based sensors.^{524, 525} However, SERS seems to be more popular technique for the development of hBN-based chemical sensors. Noteworthy, boron-nitrogen bonds of hBN can be used for further modifications or to facilitate the detection of organic molecules *via* the polarization of B-N bonds.¹⁶⁸ Moreover, the modification of the BN surface, e.g. introducing noble metal NPs, allows the enhancement in Raman the signal thus rendering such material a practical support for SERS. Moreover, BN exhibits high fluorescence quenching ability and different affinity forward single stranded DNA and double stranded DNA, allowing sensitive detection of complementary DNA fragments.⁵²⁶

Khan *et al.* have reported electrocatalytic behaviour of hBN toward the detection of dopamine.⁵²⁴ The electrode was prepared by drop-casting ethanolic solution of crystalline hexagonal boron nitride on screen-printed graphitic electrode (SPE) and occurred to determine dopamine in the presence of common interference molecules ascorbic acid (AA) and uric acid (UA). The obtained sensor exhibited limit of detection of 0.65 μ M, being comparable to other 2D-based dopamine sensors. Zhan and co-workers exploited the quenching ability of boron nitride for the DNA detection.⁵²⁶ The experiment revealed effective fluorescence decrease of labelled ss-DNA in the presence of boron nitride nanosheets (BNNS), reaching up to 85%. This sensor exhibited low limit of detection toward target DNA molecule (104 pM) and linearity in the range of 0.6-20 nM. Lin *et al.* have assembled a SERS-based sensor for rhodamine 6G (R6G) detection with high reusability.¹⁶⁸ Silver-decorated boron nitride nanosheets, prepared *via* reduction of Ag salt in the presence of BNNS, were deposited on quartz substrates using substrate transfer techniques resulting in thermal-oxidation resistant material, which facilitates removal of residue analyte. Another

SERS sensor for aromatic molecules was prepared by introduction of Au NPs onto BN nanosheets obtained by Scotch tape technique. Cai *et al.* proposed substrates tested toward R6G detection, with signals greatly enhanced in the presence of 10^{-6} M solutions of analyte.⁵²⁷ The fabricated sensor also allows the effective removal of detected molecules by thermal treatment.

MXenes-based sensors exhibit good electrochemical activity toward range of biologically important molecules like glucose,⁵²⁸ H_2O_2 ^{178, 529} and nitrites.¹⁷⁹ Wang and co-workers presented nanocomposite of TiO_2 modified with Ti_3C_2 MXene encapsulating haemoglobin with effective activity toward H_2O_2 detection.⁵²⁹ TiO_2 - Ti_3C_2 was synthesized in autoclave then the mixture consisting of suspension of nanocomposite, haemoglobin in phosphate buffered saline (PBS) and naftion were applied on GCE by simple casting method. The obtained MXene enzyme-based sensor exhibited efficient electrochemical performance allowing determination of hydrogen peroxide with low limit of detection of 14 nM and wide linear range of 0.1-380 μM . Another Ti_3C_2 -based electrochemical sensor supplied with haemoglobin to provide direct electron transfer was realized by Liu *et al.*¹⁷⁹ These sensor, due to large surface area and high conductivity, displayed good electrochemical performance allowing detection of sodium nitrite (NaNO_2) with linear range of 0.5-11800 μM and limit of detection equal to 0.12 μM . Rakhi and co-workers have used synergistic effect of Au NPs and MXenes toward glucose oxidation.⁵²⁸ Two-dimensional titanium carbide was obtained *via* exfoliation of Ti_3AlC_2 . Further sonication of MXene with Au NPs resulted in Au/MXene composite used to prepare electrode together with glucose oxidase (GOx) by casting the obtained mixture onto GCE. These sensors displayed detection of glucose up to 5.9 μM and amperometric response with linearity in the range of 0.1 mM to 18 mM.

Two-dimensional nanosheets of black phosphorous have gained attention as an analogue of other widely explored layered materials. BP has impressive opto-electronical properties but it is particularly sensible to water and oxygen, thus during the preparation of BP-based chemical sensors particular precautions have to be taken.^{530, 531} Yan and co-workers proposed nonenzymatic sensor based on solvent exfoliated BP for determination of H_2O_2 .⁵³⁰ Few-layered black phosphorous, produced via a supercritical carbon dioxide-assisted synthesis, were used for electrode preparation by casting ethanolic solution onto GCE. The electrochemical impedance was investigated in N_2 -saturated phosphate buffered saline (PBS) exhibiting linearity in the range of 1×10^{-7} M to 1×10^{-5} M with limit of detection as low as 1×10^{-7} M. Yew *et al.* used higher affinity of BP NPs toward ssDNA and prepared fluorescent

sensor for nucleic acid detection.⁵³¹ Biosensing platform provided linearity in the range of 4-4000 pM and limit of detection equal to 5.9 pM.

The performance of hBN, MXenes and BP-based materials as (bio)molecular sensors are summarized in Table 11.

Table 11. Comparison of sensing properties of hBN, MXenes and BP-based chemical sensors.

<i>Material</i>	<i>Type of sensor</i>	<i>Analyte</i>	<i>Limit of detection</i>	<i>Linearity Range</i>	<i>Ref.</i>
<i>hBN</i>	electrochemical	DA	0.65 μ M	-	524
<i>surfactant exfoliated hBN</i>			1.57 μ M	-	525
<i>BNNS</i>	fluorescent	DNA	104 pM	0.6-20 nM	526
<i>Ag-BNNS</i>	SERS	R6G	10 μ M	-	168
<i>Au/BN</i>			1 μ M	-	527
<i>TiO₂-Ti₃C₂</i>	electrochemical	H ₂ O ₂	14 nM	0.1-380 μ M	529
<i>Ti₃C₂</i>		NaNO ₂	0.12 μ M	0.5-11800 μ M	179
<i>GOx-Au-MXene</i>		Glucose	5.9 μ M	0.1-18 mM	528
<i>BP</i>		H ₂ O ₂	0.1 μ M	0.1 - 10 μ M	530
<i>BPNPs</i>	fluorescent	DNA	5.9 pM	4-4000 pM	531

List of abbreviations. DA : dopamine, BNNS : boron nitride nanosheets, SERS : Surface Enhanced Raman Scattering, R6G : rhodamine 6G, GOx : glucose oxidase, BPNPs : black phosphorus nanoparticles.

In this chapter we have discussed 2DMs and 2DM-based composites for the sensing of (bio)molecular species. A range of unique physico-chemical properties offered by 2DMs allows fabrication of various types of chemical sensors including fluorescent, electrochemical, FET and SERS. Recent literature includes investigations of most important biological molecules with particular attention to molecules relevant in medical sciences. Among various advantages of 2DMs based sensors the extremely low limit of detections (reaching the picomolar level), ultrafast detection (<1s) and wide linearity of response signals should be emphasized. Moreover, most of 2DMs based sensors exhibit high selectivity and allow determination of multiple target molecules even in the presence of common interferences. Possibility of further modifications of 2DMs allows achieving multicomponent systems offering better performances in effective determination of target analytes. Noteworthy, the spectrum of molecules detectable with presented sensors is quite narrow. Also there are a few issues that should be considered while development of new sensing

materials such as stability, reusability and biocompatibility. Electrochemical sensing seems to be one of most widespread approach applied in chemical sensing, which affirms superior electrical properties of graphene and TMD's. In case of TMD recent investigations seem to be mainly focused on the MoS₂ and WS₂ while the other representatives did not gather much attention so far. Notably the group of novel 2DMs such as hBN, MXenes and BP hold great potential in molecular sensing. Basing on few recent reports those materials constitute promising fundamentals for further investigations in terms of its sensing properties.

5 Conclusions and outlooks

Alongside the numerous extraordinary physical and chemical properties possessed by 2DMs, their high surface-to-volume ratio represents a key additional advantage for applications in sensing. The all-surface-and-volume-free character guarantees a highest sensitivity to any changes of 2DMs' environment, which is an absolute prerequisite for the development of high-performance sensing devices. However, the use of pristine, defect free sheets may result into weak and unspecific interactions with the chosen analytes. Defects hold a pivotal role in 2DMs: they allow the modulation of the (opto)electronic properties and in particular the engineering of the band gap, which plays a crucial role in improving the sensitivity of the device. Moreover, the presence of dangling groups on the surface can promote analyte absorption, consequently yielding an improvement in the sensing performances. Furthermore, the presence of defects highly increases chemical reactivity of 2DMs enabling further chemical functionalization of the reactive sites, which may result into high specificity, enhanced sensitivity and low detection limit towards targeted analyte. On the other hand, a lack of control during defect engineering can result in irreproducible properties of 2DMs. In particular, the increased reactivity and the presence of dangling bonds with different chemical structures on the basal plane (or edges) of 2DM may result in an enhanced, yet, unselective binding of the analytes ultimately leading in poor selectivity, long response or even incomplete recovery time of the sensor.

One of the greatest challenges to be tackled in the near future is the controlled introduction of *ad hoc* defects in the 2DMs with a precise spatial resolution and concentration. Ideally, the chemical composition of the defects has to be complemented by a chemical functionalization of the reactive sites with the receptor of the analyte of choice. Such a strategy needs to be pursued by mastering principles of molecular recognitions, e.g. *via* tailoring of

supramolecular interactions, which may result not only in ultrasensitive 2DMs but also in superior selectivity and response time - making the actual step towards industrial products. Additionally, chemical programmability, through the generation of multicomponent assemblies based on 2DMs, may enable the future development of multianalyte sensing. Although optical sensors based on the change of the photoluminescent properties of the 2DMs upon interaction with the analytes have been widely and successfully employed for both metals and biologically relevant molecules, electrochemical sensors seem to be the most promising since they provide a direct electric response without the need of employing a transducer. By taking advantage not only of the superior electrical properties of graphene and other 2DMs, but also of their excellent mechanical properties such as robustness and flexibility, 2DMs based sensors can be supported on flexible foils, for smart wearable technologies for monitoring the human/environment interface. Interestingly, the recent demonstration the compatibility of 2DMs with “on-the-skin” applications opens the door towards their use for the detection of clinically relevant biomarkers. Such 2DM-based tattoos may therefore find application as dry on-the-skin sensors characterized by low susceptibility to motion artefacts, which is one of the biggest drawbacks of conventional dry sensors and electrodes for point-of-care and health monitoring.

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