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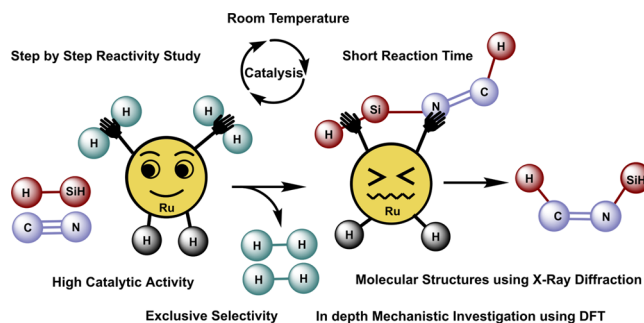
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Mapping Out the Role of σ -Silane Complexes in the Ruthenium-Catalyzed Hydrosilylation of Nitriles

Aswin Chandran, Raphael Ravel-Massol, Laure Vendier, John M. Slattery,* Jason M. Lynam,* and Mary Grellier*

ABSTRACT: A combined synthetic, mechanistic, and computational study is reported, which provides unique insight into the role of σ -silane complexes in the catalytic hydrosilylation of nitriles. A novel, highly efficient, highly active, and regioselective catalytic monohydrosilylation of aromatic nitriles with secondary silanes using a ruthenium dihydrogen catalyst is reported along with a novel mechanism for hydrosilylation of nitriles. Investigations into the mechanism of this transformation have revealed the influence of σ -Si-H complexes in fine-tuning the selectivity of this hydrosilylation reaction. Displacement of the dihydrogen ligand on the ruthenium precatalyst, ruthenium bis-(dihydrogen) complex $[\text{RuH}_2(\eta^2\text{-H}_2)_2(\text{PCy}_3)_2]$, **1**, by diphenylsilane leads to the formation of new ruthenium σ -Si-H complexes, $[\text{RuH}_2(\eta^2\text{-H}_2)(\eta^2\text{-HSiHPh}_2)(\text{PCy}_3)_2]$, **2**, and $[\text{RuH}_2(\eta^3\text{-H}_2\text{SiPh}_2)(\text{PCy}_3)_2]$, **3**. Complex **3** reacts readily with benzonitrile leading to hydrosilylation of the nitrile and coordination of the silylimine formed to the ruthenium as a σ -H-Si-N-silylimine complex, $[\text{RuH}_2(\eta^2\text{-HSiPh}_2\text{NCHPh})(\text{PCy}_3)_2]$ (**4**). This systematic investigation of this reactivity led to the discovery of the first direct evidence of an *N*-silylimine-coordinated ruthenium complex and its involvement in a catalytic hydrosilylation reaction. This led to the discovery of a catalytic protocol for the efficient and selective coupling of secondary silanes with a range of nitriles using **1** as the catalyst. It is proposed that complexes **3** and **4** are key intermediates on the catalytic reaction coordinate, which leads to hydrosilylation of the nitrile. This is supported by DFT calculations along with the observation that **3** and **4** are catalytically active. The Si-N bond formation was found to proceed via direct attack of the nitrile at the silicon atom in **3**. Through carefully chosen structural studies and tests of the new ruthenium complexes, along with DFT calculations, the mechanism of the catalytic hydrosilylation of nitriles has been successfully explained.



INTRODUCTION

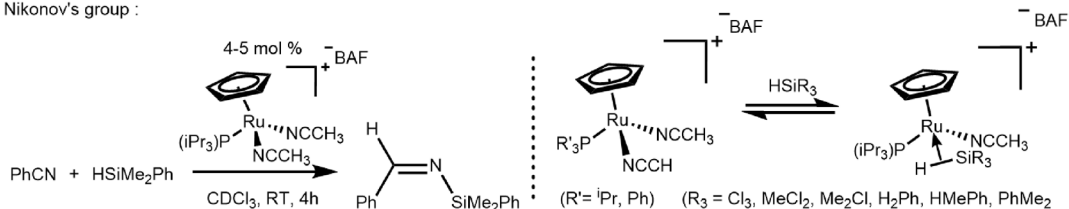
The significance of hydrosilylation of C-N triple bonds in organonitriles is well-established, as the resulting organosilicon products have a wide range of applications, particularly in organic synthesis.¹⁻⁵ The hydrosilylated products formed, such as disilylamines, are used as precursors for the synthesis of polymers containing Si-N bonds, such as polysilazanes and polysilsesquiazanes, which are used as precursors for ceramic applications.^{6,7} *N*-Silylamines are synthons for various catalytic transformations,⁸ including C-N bond-forming cross-coupling reactions;⁹ meanwhile, *N*-silylimines are well-established for their use in enantioselective synthesis.¹⁰⁻¹⁵ However, the catalytic monoaddition of a Si-H bond to a C-N triple bond continues to be a synthetic challenge, which makes this transformation a key target for new catalysts.¹⁶⁻¹⁸ Controlling the selectivity of this transformation is particularly challenging as the initially formed *N*-silylimines, R₃Si-N=CHR, may readily undergo a second addition to form the double-addition product disilylamines, (R₃Si)₂NCH₂R. However, a small number of metal¹⁹⁻²³ and nonmetal-based^{24,25} catalysts for

monohydrosilylation of nitriles to form *N*-silylimines selectively have begun to emerge recently. The interaction of hydrosilanes with transition metals has been well-documented in the series of reviews by Corey.²⁶⁻²⁸ The catalytic monohydrosilylation of organonitriles with monohydrosilanes to synthesize *N*-silylimines selectively was achieved by Nikonov and co-workers using a cationic ruthenium complex $[\text{CpRu}(\text{NCMe})_2(\text{P}^i\text{Pr}_3)][\text{B}(\text{C}_6\text{F}_5)_4]$ as a precatalyst with almost quantitative yields after 0.3 to 48 h at ambient temperature with a 4-5 mol % catalytic loading (Scheme 1a).¹⁶ A mechanism was proposed involving the formation of a ruthenium σ -Si-H silane complex, $[\text{CpRu}(\text{NCMe})(\text{P}^i\text{Pr}_3)(\eta^2\text{-HSiR}_3)]^+$, followed by nucleophilic abstraction of the silylium

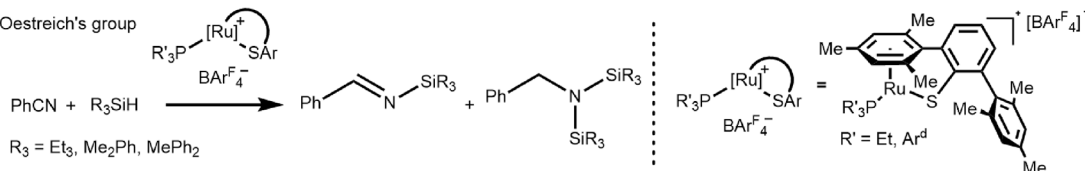
Scheme 1. (a) Catalytic Chemoselective Mono- or Bis-hydrosilylation of Nitriles Using a Ruthenium Catalyst and Synthesis of a Ruthenium σ -Si-H Silane Complex Implicated in the Reaction, (b) Catalytic Hydrosilylation of Nitriles Using a Ruthenium-Sulfur Complex, (c) Previous Results for the Reactivity of Complex 1 with Secondary Silanes and Benzonitrile, and (d) New Insights and Novel Catalysis from This Work

Previous work

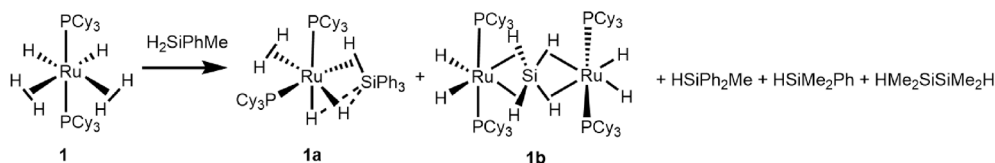
(a) Nikonov's group :



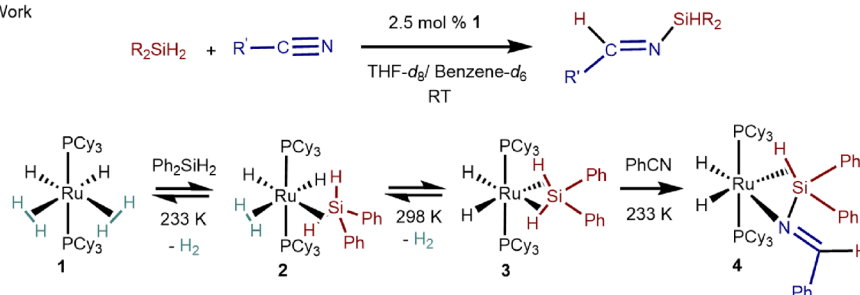
(b) Oestreich's group



(c) S. Etienne and M. Grellier's group :



(d) This Work



- Role of σ -complexes in hydrosilylation of nitriles
- Isolation and catalytically active complexes 3 and 4
- Mechanistic studies using DFT calculations

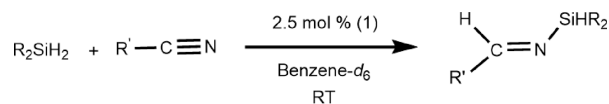
ion by one of the nitrile ligands to form a labile $[\text{CpRuH}(\text{NCMe})(\text{P}^i\text{Pr}_3)]$ complex. Oestreich et al. reported a different cationic ruthenium catalyst²¹ (Scheme 1b), which effectively led to the formation of mono- and bis-hydrosilylated products depending upon the specific hydrosilane used. Oestreich et al. have studied the mechanism of cooperative Si-H bond activation using the Ru-S catalyst.²⁹ However, a fundamental and detailed study of the interactions between the silane, the nitrile functional group, and the ruthenium center along the pathway leading to Si-N and C-H bond formation needs to be investigated for a selective monohydrosilylation of nitrile using secondary silane. Here, we address this gap in knowledge by utilizing the ruthenium bis-dihydrogen dihydride complex, $[\text{RuH}_2(\eta^2\text{-H}_2)_2(\text{PCy}_3)_2]$ (**1**), which is well-established to activate silanes and forms a number of well-characterized σ -complexes.³⁰⁻³⁶ It was reasoned that **1** might be a highly active catalyst for the hydrosilylation of nitriles, as it is a well-established reagent for the activation of the Si-H bonds

(Scheme 1c).^{32,35} In this work, the catalytic activity of complex **1** in the nitrile hydrosilylation is reported. New ruthenium σ -Si-H complexes of fundamental interest have been characterized and isolated, and their role in the catalytic reaction has been identified. Combining these observations with DFT calculations enabled us to propose a reaction mechanism for the selective monohydrosilylation of nitriles using secondary silanes.

RESULTS AND DISCUSSION

Catalytic Performance of $[\text{RuH}_2(\eta^2\text{-H}_2)_2(\text{PCy}_3)_2]$ (1**) for the Hydrosilylation of Nitriles.** The fundamental premise of this work is that the established role of **1** in the coordination and activation of element-hydrogen bonds would (1) entail that it would be a catalyst for the hydrosilylation of nitriles and (2) enable mechanistic insight into the catalytic cycle, as the key Si-H σ -complexes were expected to be sufficiently stable to enable their role in the Si-N and C-H bond formation

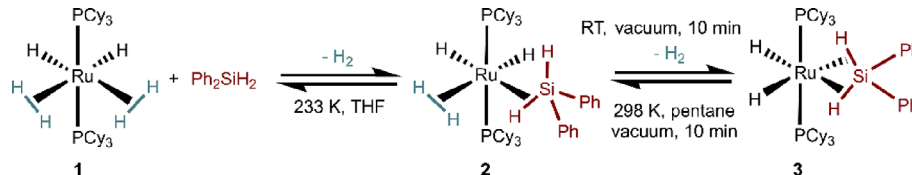
Table 1. Catalytic Hydrosilylation of Nitriles with Secondary Silanes^d



Entry	Nitrile	Silane	TON	TOF [h ⁻¹] ^a	Conversion ^b (min)	Yield ^c (%)	Product
1	PhCN	Ph ₂ SiH ₂	40	185 (100%)	>99% (13)	96%	
2	PhCN	Et ₂ SiH ₂	40	171 (100%)	>99% (14)	97%	
3	PhCN	PhMeSiH ₂	40	400 (100%)	>99% (6)	87%	
4		Ph ₂ SiH ₂	40	240 (100%)	>99% (10)	89%	
5		Ph ₂ SiH ₂	40	240 (100%)	>99% (10)	96%	
6		Ph ₂ SiH ₂	38	53 (44.5 %)	>99% (150)	85%	
7		Ph ₂ SiH ₂	40	200 (100%)	>99% (12)	86%	
8		Ph ₂ SiH ₂	40	160 (100%)	>99% (15)	99%	
9		Ph ₂ SiH ₂	40	160 (100%)	>99% (15)	89%	
10	PhCN	2 Ph ₂ SiH ₂	40	160 (100%)	>99% (15)	100%	 No further reaction after heating 60 °C for 1 day.
11	CH ₃ CN	Ph ₂ SiH ₂	0	0	0	0	No catalytic reaction

^aCalculated at the conversion shown in parentheses. ^bCalculated from ¹H NMR data. ^cYield is determined by ¹H NMR. ^dGeneral procedure: 2.5 mol % [RuH₂(η²-H₂)₂(PCy₃)₂] was added to a solution of nitrile and silane in C₆D₆.

Scheme 2. Reversible Coordination of Ph₂SiH₂ to Complex 1 to Form Complexes 2 and 3



steps to be identified. Therefore, the ability of compound **1** to catalyze the desired reaction was investigated. Catalytic hydrosilylation of nitriles with a series of secondary silanes and nitriles was observed with 2.5 mol % **1** (Table 1). The catalyst was found to be highly active and efficient, achieving nearly complete conversion and an excellent yield at room temperature. The complete conversion was typically observed in 15 min with a maximum TOF of 400 h⁻¹ and 99% yield (Table 1, entry 3). In all cases, only the monoaddition of one of the Si–H bonds was observed: even with 2 equiv of diphenylsilane (or a second equivalent was added when the first equivalent was consumed), no evidence for double silane addition giving the corresponding amine was obtained (Table 1, entry 10).

Complex **1** was found to be at least as efficient for the catalytic monohydrosilylation of nitriles when compared to previously reported ruthenium complexes, but in many cases, the performance is superior. For example, **1** exhibits turnover frequencies of 240 (at 50% conversion) and 185 h⁻¹ (at 100% conversion) for the catalytic reaction between PhCN and Ph₂SiH₂, while the catalyst [CpRu(NCMe)₂(PiPr₃)] [B(C₆F₅)₄] shows a turnover frequency of only 6.3 (at 50%) for the similar monohydrosilylation of PhCN using PhMe₂SiH at room temperature.¹⁶ Compared to the Ru–S catalyst,²¹ the reaction time under mild conditions is significantly reduced from 18 to 1 h.

Complex **1** was able to effectively catalyze the hydrosilylation of a number of aryl nitriles (entries 4–9, Table 1) but was less effective with alkyl nitriles. No catalytic activity was observed in the case of coupling acetonitrile with diphenylsilane under the same conditions (Table 1, entry 11). No compound resulting from a mono- or bis-hydrosilylation was detected. This chemoselectivity for aryl nitrile was exploited for the hydrosilylation of 4-cyanophenylacetonitrile (Table 1, entry 9). Only the reduction of the cyano fragment attached to the aromatic ring was hydrosilylated in a good yield (89%). Different monohydrosilanes such as Ph₃SiH, PhMe₂SiH, and Et₃SiH were also tested with PhCN in the same catalytic reaction conditions, and no hydrosilylated product formation was observed.

Formation of Novel Ruthenium σ -Silane Complexes 2 and 3. With the excellent catalytic activity of **1** observed and the substrate scope explored, selected stoichiometric reactions were conducted to understand the mechanism of the reaction. To achieve this, the interactions between **1** and the individual components of the catalytic reaction were investigated. Benzonitrile was chosen as the model aromatic nitrile, and diphenylsilane was chosen as the model secondary silane for stoichiometric reactions. The reactivity of **1** with nitriles has been documented in a patent.³⁷ The reactivity of benzonitrile with analogous complex [RuH₂(η²-H₂)₂(PCy₃)₂], **1-PCy₃**, has been investigated previously and, depending on the

stoichiometry, gives products that ultimately derive from N coordination of PhCN to the metal and may also result in C–H bond activation of the phenyl group.³⁸

It is well-established that the ruthenium bis-dihydrogen complex, **1**, may act as a precursor to ruthenium σ -Si–H complexes.^{32,33} However, a comprehensive study of the coordination of diphenylsilane at low temperatures has not been conducted. A previous study of the coordination of a dihydrosilane, H₂SiPhMe, to **1** at room temperature revealed a redistribution reaction of the silane and formation of complexes **1a** and **1b** along with a mixture of different silanes (Scheme 1c). To understand the intimate interaction of diphenylsilane with complex **1**, low-temperature conditions were employed. These conditions enabled the spectroscopic identification of two new ruthenium complexes. The addition of diphenylsilane to a THF solution of **1** at 233 K followed by slow warming to room temperature resulted in the formation of σ -Si–H complexes [RuH₂(η²-H₂)(η²-HSiHPh₂)(PCy₃)₂], **2**, and [RuH₂(η³-H₂SiPh₂)(PCy₃)₂], **3** (Scheme 2).

To obtain a clearer understanding of the mechanism of silane coordination to **1**, the reaction between **1** and Ph₂SiH₂ was performed in an NMR tube at 223 K just prior to introduction into the spectrometer, preset at the same temperature. The resulting ³¹P{¹H} NMR spectrum showed a single peak at 56.1 ppm, which was assigned to a new species, [RuH₂(η²-H₂)(η²-HSiHPh₂)(PCy₃)₂], **2**, that possessed a single phosphorus environment. The ¹H NMR spectrum exhibited a triplet resonance at 5.87 ppm (³J_{PH} = 5 Hz) with two silicon satellites (¹J_{SiH} = 194 Hz), which, based on its chemical shift and Si–H coupling, was assigned to a Si–H group that was not coordinated to the ruthenium. A broad signal at –8.83 ppm that integrated to five hydrogens was assigned to the ruthenium-coordinated hydrogen atoms, as a similar resonance was observed in the known silane complex [RuH₂(H₂)(η²-HSiPh₃)(PCy₃)₂] (–8.3 ppm (br)),³⁹ in which the ruthenium hydrides, the dihydrogen ligand, and Si–H are all undergoing rapid exchange. The same assignment is therefore made in this case. In the ²⁹Si DEPT spectrum, a singlet signal was observed at –9.9 ppm. This resonance exhibited a correlation with the hydride signal and the free Si–H signal in a ¹H–²⁹Si HMQC experiment (SI, Figures S16 and S17). The presence of a dihydrogen ligand in complex **2** is supported by the minimum T_{1,min} value (67 ms at 283 K and 600 MHz). The optimized structure of **2** by DFT also confirms the η²-H₂ coordination (see SI, Section 10).

It was not possible to isolate complex **2** as it subsequently converted to complex **3**; however, the coordination of diphenylsilane to **1** was shown to be reversible. The reaction of isolated complex **3** in THF-*d*₈ solution with 1.5 bar H₂ pressure resulted in the sequential formation of complex **2** followed by precursor bis-dihydrogen complex **1** (Scheme 2). The reaction was monitored by ¹H and ³¹P{¹H} NMR

spectroscopy, and within 30 min at room temperature, the disappearance of signals corresponding to **3** and the appearance of signals corresponding to **2** and **1** were complete. After 1 day, further conversion to **1** was observed based on the integration values of the signals in $^{31}\text{P}\{^1\text{H}\}$ and ^1H NMR (SI, Figures S44 and S45).

The formation of **3**, $[\text{RuH}_2(\eta^3\text{-H}_2\text{SiPh}_2)(\text{PCy}_3)_2]$, may be rationalized as a substitution for the dihydrogen ligand in complex **2** by the pendant Si–H bond. In complex **3**, the silane may be considered as a bis- σ -Si–H complex in which both Si–H hydrogen atoms are coordinated in an η^2 fashion to the metal. The conversion from **2** to **3** was observed at room temperature, and the use of a solvent in which complex **3** is sparingly soluble, along with a reduced pressure by applying a vacuum using a Schenk line, helped to accelerate the conversion. The removal of dihydrogen led to a shift in equilibrium toward **3**. The complete conversion was observed after 7 h in THF; meanwhile, using pentane as a solvent and applying a vacuum to remove the dihydrogen released reduced the reaction time to under 10 min at room temperature. This technique led to the successful isolation of **3** with a yield of 92%.

The molecular structure of **3** was determined by NMR spectroscopy and X-ray crystallography, along with optimization by using DFT modeling. The ^1H NMR spectrum of **3** in THF- d_8 exhibited a broad singlet resonance at -8.36 ppm (4H, RuH) at 298 K due to fast exchange between the hydride ligands. A singlet at 67.9 ppm was also observed in the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum. Decreasing the temperature led to decoalescence of the high-field signals at a temperature below 273 K. At 193 K, the ^1H NMR spectrum exhibited a broad singlet around -5.8 ppm and a triplet at -12.33 ppm ($^2J_{\text{P-H}} = 25.8$ Hz, 2H, RuH) (SI, Figures S23 and S31). The $^1\text{H}\{^{31}\text{P}\}$ NMR spectrum at 193 K gives a sharpened signal at -5.82 ppm, leading to the observation of the silicon satellites with a $^1J_{\text{Si-H}}$ of 122 Hz. In this experiment, the signal at -12.33 was then simplified into a singlet with no detectable silicon satellites. These high-field resonances exhibited a mutual correlation in the ^1H – ^1H COSY experiment and in the ^1H – ^{31}P correlation experiment; both were coupled to the resonance in the $^{31}\text{P}\{^1\text{H}\}$ spectrum at 67.9 ppm. These observations are consistent with the proposed formation of **3**: the resonance at -5.82 ppm was assigned to the two hydrogen atoms involved in the Si–H interaction, with the ruthenium hydride resonance at -12.33 ppm being assigned to a classical dihydride. The T_1 measurements recorded (between 694 and 356 ms at 600 MHz) with no minimum value over a range of temperatures (193 to 253 K) for the signal at -12.33 ppm are in agreement with classical dihydride coordination in complex **3** (SI, Table S1).

Although direct detection of a ^{29}Si resonance was not possible at any of the temperatures employed, ^1H – ^{29}Si HMQC and 1D HSQC NMR (193 K) experiments demonstrated the presence of a resonance in the ^{29}Si NMR spectrum at 120.8 ppm ($^1J_{\text{Si-H}} = 122.7$ Hz). A similar downfield ^{29}Si chemical shift was obtained for ruthenium bis- σ -Si–H complexes reported by Lipke and Tilley $[\text{PhBP}^{\text{Ph}}_3]\text{RuH}(\eta^3\text{-H}_2\text{SiRR}')$ ($\text{RR}' = \text{PhMe}$ or Ph_2 , $(\text{PhBP}^{\text{Ph}}_3) = \text{tris}(\text{diphenylphosphinomethyl})\text{borato}$ ligand) at 154 ($^1J_{\text{Si-H}} = 65$ Hz) and 141 ppm ($^1J_{\text{Si-H}} = 68$ Hz).⁴⁰ The ^{29}Si chemical shift is in agreement with the proposal that **3** contains a bis- σ -Si–H ligand, rather than a silylene, as the latter is usually observed at >200 ppm.^{40–42} Meanwhile, the higher $J_{\text{Si-H}}$

coupling constant in **3** suggests that the Si–H bonds are weakly activated with a stronger Si–H interaction compared to the examples reported by Lipke and Tilley. However, the Si–H coupling constant in **3** is within the range typically observed for a σ -Si–H interaction.³³ The characteristic chemical shift in the region for a σ -Si–H bond, along with a stronger Si–H interaction, ruled out the possibility of the complete oxidative addition and identified the nonclassical Ru–H–Si interaction.

Crystals of **3** suitable for study by single-crystal X-ray diffraction were obtained by the slow cooling of a pentane-THF solution of **3** to 233 K. The ORTEP view of the molecular structure obtained by X-ray diffraction analysis of **3** is shown in Figure 1. The structure revealed that the complex

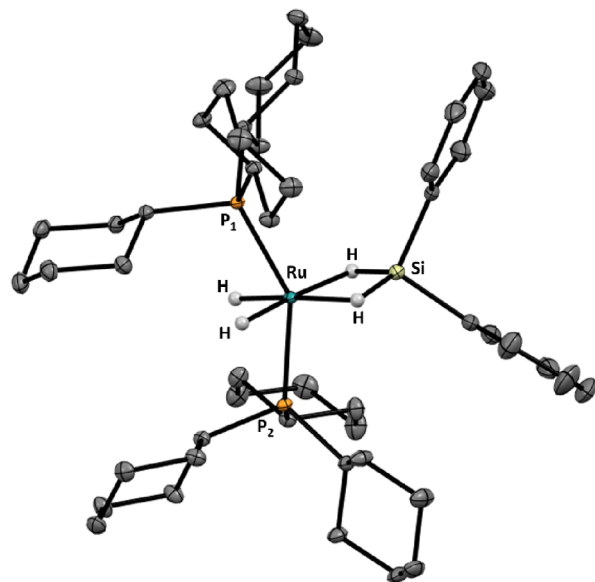


Figure 1. X-ray molecular structure of complex **3**, $[\text{RuH}_2(\eta^3\text{-H}_2\text{SiPh}_2)(\text{PCy}_3)_2]$, drawn at 50% probability as ellipsoids (hydrogen atoms have been removed for clarity with the exception of those around ruthenium).

exhibits two phosphine groups in a mutually trans geometry and two classical ruthenium hydrides and two η^2 -Si–H hydrides in a plane perpendicular to the phosphine ligands, leading to a pseudo-octahedral geometry around the ruthenium center. The hydrogen atoms are also located around the ruthenium in a plane perpendicular to the phosphine ligands. The ruthenium–silicon distance was found to be 2.2466 (10) Å; meanwhile, the shortest distances from the ruthenium atom to the two Si–H hydrogen atoms were 1.83 (3) and 1.79 (3) Å, respectively, which are comparable to a similar magnitude as reported in the complexes $[(\text{PhBP}^{\text{Ph}}_3)\text{RuH}(\eta^3\text{-H}_2\text{SiRR}')] (\text{RR}' = \text{PhMe}$ or Ph_2 (1.73 (4) and 1.76 (4) Å)).⁴⁰ The Si–H bonds were found to be 1.53 (3) and 1.49 (3) Å, compared to free Si–H bonds of silane (1.48 Å), suggesting that the Si–H bonds are weakly activated. This structural arrangement is consistent with the data from the NMR spectra, notably at low temperatures, with the observation of two different resonances in the high-field region: one for the classical dihydrides and one for the σ -Si–H bonds. The geometric environment around the ruthenium atom is similar to an analogous ruthenium borane complex $[\text{RuH}_2(\eta^2\text{-}\eta^2\text{-H}_2\text{BMes})(\text{PCy}_3)_2]$.⁴³ The similarity in the geometry is also manifested in the characteristic ^1H NMR signals of this borane complex. The ruthenium hydrides were

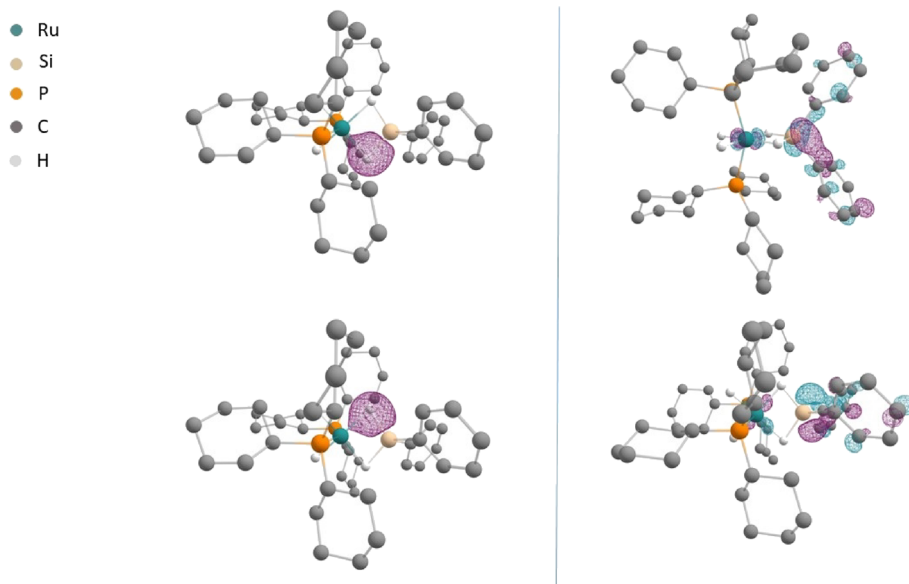


Figure 2. Left: pNLMO representation of the two $\text{Ru}(\eta^2\text{-SiH})$ bonds. Right: LUMO of compound **3** (two perpendicular views). Hydrogen atoms attached to carbon atoms have been omitted for clarity.

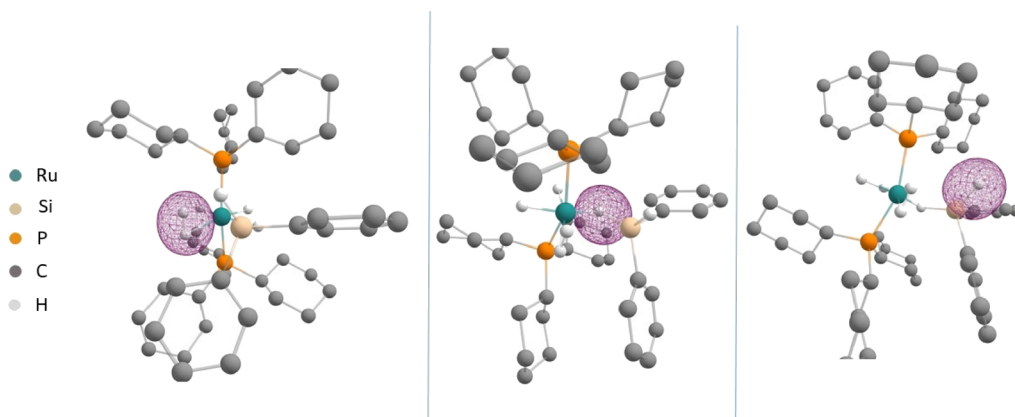


Figure 3. pNLMO representation of (left) $\text{Ru}(\eta^2\text{-H}_2)$, (middle) $\text{Ru}(\eta^2\text{-SiH})$, and (right) uncoordinated Si-H , complex **2**. Hydrogen atoms attached to carbon atoms are omitted for clarity.

observed at -5.90 (s, 2H) and -11.05 ppm (t, $^2J_{\text{P-H}} = 25.2$ Hz),⁴³ which are comparable to complex **3**.

DFT calculations were used to optimize the geometry of **3**, and particular attention was given to locating the hydrogen atoms around ruthenium. The bond lengths and angles are in good agreement with the X-ray molecular structure. For instance, the calculated distance between ruthenium and silicon atoms was determined to be 2.227 Å, which is in close agreement with the experimentally obtained measurement of 2.2466 (10) Å. A similar correlation was observed for Ru–P bond lengths, with calculated values of 2.304 and 2.311 Å aligning closely with XRD-obtained values of 2.3275 (9) and 2.3338 (8) Å, respectively. The calculated distances between the ruthenium atom and the two Si–H hydrogen atoms were 1.83 Å, and the distance between the silicon atom and two Si–H hydrogen atoms was found to be 1.65 Å each. Characteristic bond lengths obtained by X-ray diffraction in comparison with the calculated values are presented in SI, Table S2.

In order to understand the coordination of Ph_2SiH_2 to the ruthenium center, an NBO analysis was performed on complex **3**. While the σ -donating NBOs have nearly the same

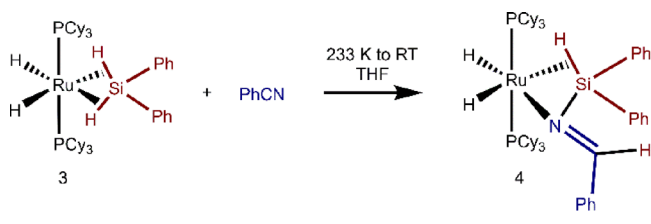
population in **1** (1.70 e, σ (H_2)) and **3** (1.73 e, σ Si–H), the π back-bonding in **3** is slightly smaller (0.09 e, σ^* Si–H) than that in **1** (0.14 e, σ^* (H_2)). The NLMO's are also consistent with a bis- σ -Si–H description (see the pLMNO representation in Figure 2).

The NBO analysis of complex **2** confirms the σ -Si–H bond coordination mode. The donating NBOs have nearly the same population in (1.64 e, σ (H_2)) and (1.61 e, σ Si–H) (for comparison, 1.97 e, uncoordinated Si–H), but the π back-bonding is bigger (0.25 e, σ^* Si–H; 0.16 e, σ^* (H_2)) than that in **3**. The NLMO's are also consistent with a σ -Si–H, σ - H_2 bonding description (see the pLMNO representation in Figure 3).

Si–N Bond Formation from a σ -Si–H Complex and PhCN. Having established the nature of the interaction between **1** and Ph_2SiH_2 , our attention turned toward developing an understanding of the steps involved in Si–N and C–H bond formation. To achieve this, the stoichiometric reaction of complex **3** with 1 equiv of benzonitrile was explored. Specific attention was given to the potential addition of the activated Si–H bond to the benzonitrile. It was found

that a new ruthenium complex, **4**, is formed, which contains an *N*-silylimine ligand resulting from the selective addition of only one Si–H bond of diphenylsilane to the nitrile group (Scheme 3).

Scheme 3. Reactivity of Complex 3 with Benzonitrile



Addition of PhCN to a THF solution of **3** at 233 K followed by warming to room temperature afforded a new ruthenium complex with a single resonance in the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum at 66.1 ppm: this was assigned to complex **4**. The high-field region of the ^1H NMR spectrum of **4** showed three hydride signals. At room temperature, these resonances appear as broad signals; however, on cooling to 273 K, the signals sharpened, allowing for important structural information to be obtained. A singlet at -5.86 ppm with satellite peaks of $^1J_{\text{Si-H}} = 130.5$ Hz was assigned to a σ -bound Si–H group, whereas triplet resonances at -10.26 ($^2J_{\text{P-H}} = 25.4$ Hz) and -14.31 ppm ($^2J_{\text{P-H}} = 25.8$ Hz) were assigned to two inequivalent ruthenium hydride ligands. Although the coupling between phosphorus and silicon-bound hydrogen atoms was not observed, likely due to the low coupling constant values and the broadness of the signals, the correlation between them was established by a ^1H – ^{31}P HMQC experiment. The $^{31}\text{P}\{^1\text{H}\}$ signal at 66.1 ppm (s) became a quartet when the resonances for the cyclohexyl hydrogen atoms were selectively decoupled, confirming the presence of three hydrogen atoms bound to ruthenium. Interestingly, the ^1H – ^{29}Si HMQC experiment found that the $^1J_{\text{Si-H}}$ coupling constant for the ^{29}Si signal at -43.6 ppm was 128.8 Hz, which agrees with the weakly activated σ -Si–H bond formulation. A resonance at 10.1 ppm in the ^1H NMR spectrum integrating to one hydrogen and the single bond correlation to a signal at 161.7 ppm observed in ^1H – ^{13}C HSQC NMR corresponding to a CH proved the presence of the imine functional group in **4**.

The molecular structure of **4** was confirmed by single-crystal X-ray diffraction (Figure 4), which demonstrated the coordination of *N*-silylimine to ruthenium via an η^2 -Si–H bond and nitrogen atom. The ruthenium is in a pseudo-octahedral environment and is located in a plane along with Si, N, and hydrogen atoms. The distance between Ru and Si atoms was found to be 2.3503(9) Å, which is elongated compared to the precursor complex **3**, as expected. The structure provided the following bond lengths Si–N 1.740 (4), Ru–N 2.198 (3), Ru–H 1.5085 (3), 1.6557 (3), Ru–H_{Si} 1.7183 (3), and N–C 1.285 (6) Å and bond angles P–Ru–P 146.94 (3)° and Si–N–C 133.5 (3)°. The formation of new Si–N and C–H bonds and the reduction of nitrile bonds to imine bonds were confirmed with these values. The comparison of selected experimental and computational bond lengths from a DFT-optimized structure is given in SI, Table S2.

To the best of our knowledge, the formation of **4** represents the first report of an *N*-silylimine-coordinated ruthenium complex. Hashimoto and co-workers studied similar reactivity

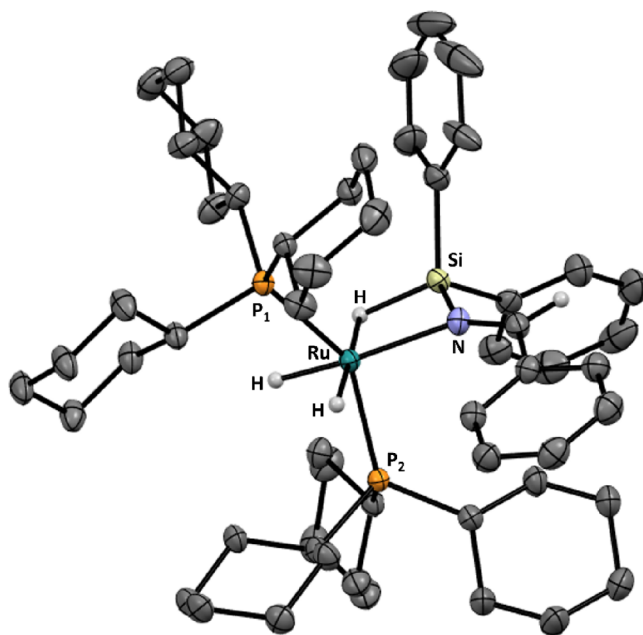


Figure 4. X-ray molecular structure of complex **4** drawn at 50% probability as ellipsoids (hydrogen atoms have been removed for clarity with the exception of those around ruthenium and NCH).

of a ruthenium complex $[\{\text{Ru}(\text{CO})_2(\text{SiTol}_2\text{H})\}_2(\mu\text{-dppm})(\mu\text{-}\eta^2\text{-}\eta^2\text{-H}_2\text{SiTol}_2)]$ with benzonitrile, and they obtained a ruthenium μ -iminosilyl complex, $[\text{Ru}_2(\text{CO})_4(\mu\text{-dppm})(\mu\text{-SiTol}_2)(\mu\text{-R}'\text{CH}=\text{NSiTol}_2)]$ stoichiometrically.⁴⁴

Further Mechanistic Study and Computational Chemistry.

The stoichiometric reactions gave insights into the mechanism of coordination of diphenylsilane as well as benzonitrile to catalyst **1**; however, in order to demonstrate the relevance of these observations to the catalytic cycle, the activity of the isolated ruthenium σ -complexes, the silane-activated complex, **3**, and the *N*-silylimine-coordinated complex, **4**, was assessed. Reactions were performed with isolated complexes **3** and **4** in the same manner as described in Table 1, and similar catalytic activity was observed (SI, Section 9.4). Specifically, when 1–2 mol % catalyst (**3** or **4**) was added to a mixture of diphenylsilane and benzonitrile, complete conversion to the corresponding *N*-silylimine, *N*-(diphenylsilyl)-1-phenylmethanimine (PhCHNSiHPh₂) was obtained exclusively at room temperature within 5 h. These experiments demonstrate that complexes **3** and **4** are catalytically competent and are potential intermediates in the catalytic reaction coordinate. Meanwhile, the addition of an excess of diphenylsilane after 1 equiv of benzonitrile led to a mixture of ruthenium diphosphine complexes, including complexes **1**, **2**, **3**, and unknown organometallic complexes. This suggests that initial diphenylsilane coordination to the metal and nitrile binding are needed to achieve the selectivity toward the *N*-silylimine. The reaction of complex **4** with 1 equiv of diphenylsilane at 233 K and warming to room temperature led to a mixture involving the ruthenium σ -silane complexes **2** and **3**, proving the formation of **4** from **2** reversibly in the presence of an excess of silane (SI, Section 9.5).

These stoichiometric and catalytic experiments successfully established (1) the role of σ -silane complexes and (2) their relevance in catalytic activity in the catalytic hydrosilylation of nitrile reaction. Computational chemistry techniques were used to explore the mechanism of the catalytic reaction. The

Scheme 4. DFT-Calculated Steps of the Catalytic Hydrosilylation Process at 298 K, Involving the Nitrile Coordination to the Silicon Center

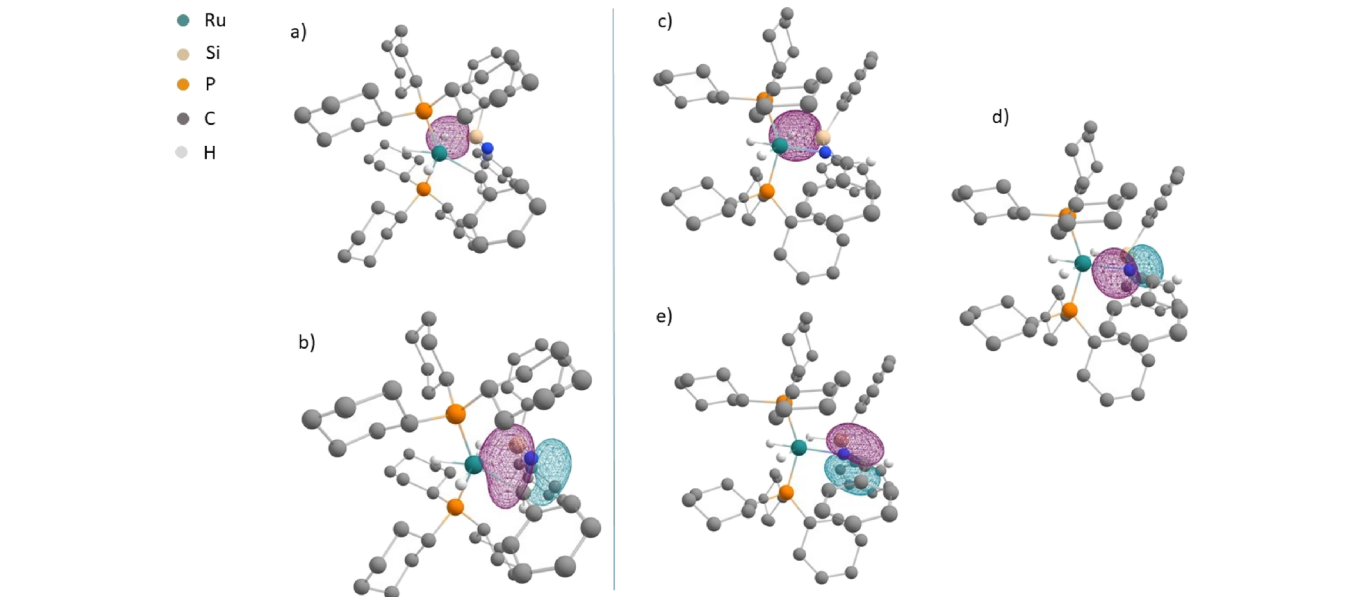
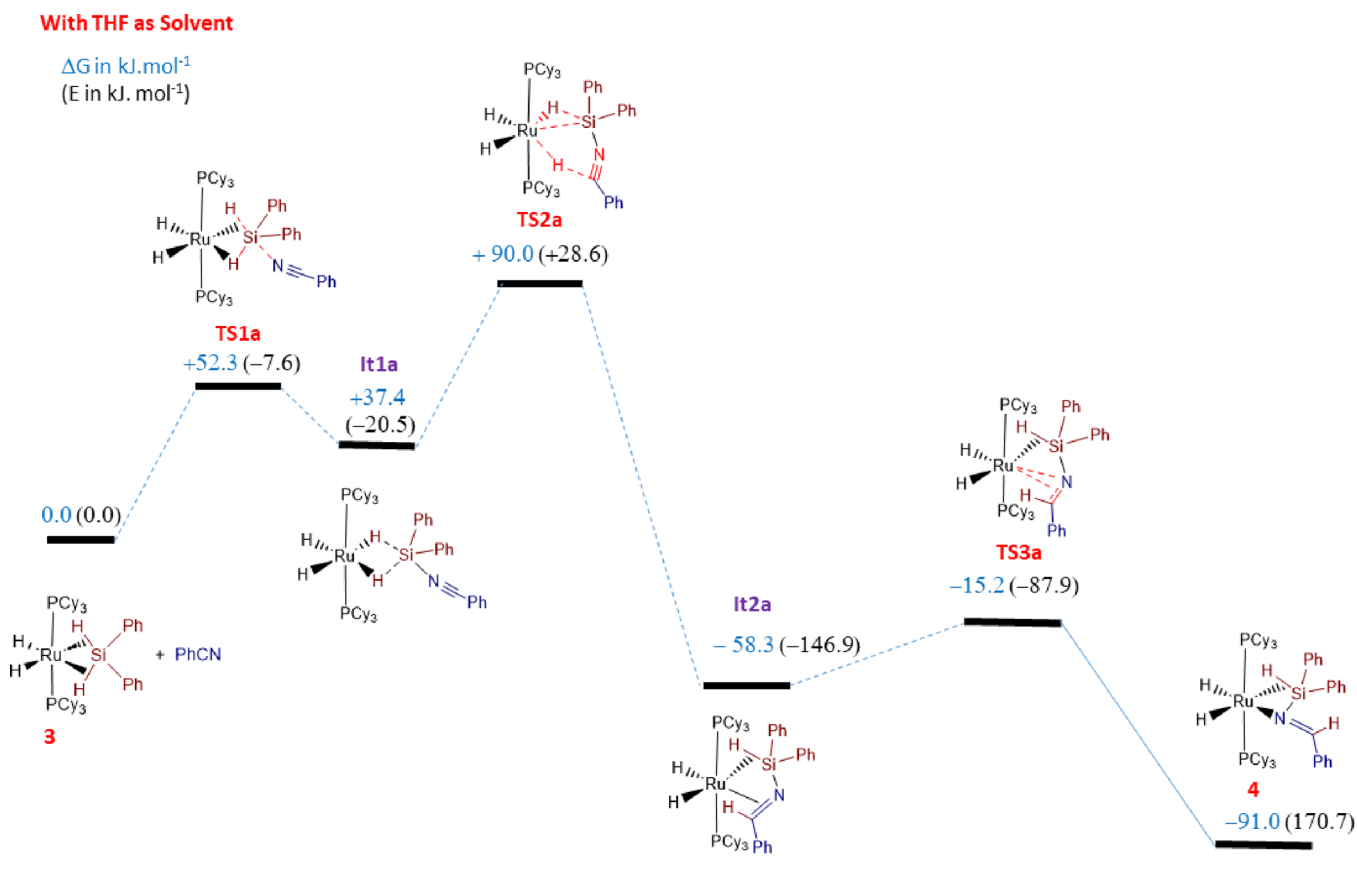


Figure 5. pNLMO representation at the left: **It2a** intermediate (a) $\text{Ru}(\eta^2\text{-SiH})$ and (b) $\text{Ru}(\eta^2\text{-C}\equiv\text{N})$; right: complex **4** (c) $\text{Ru}(\eta^2\text{-SiH})$, (d) $\text{Ru}(\kappa\text{N-N}=\text{C})$ and (e) $\text{N}=\text{C}$ bond. hydrogen atoms attached to carbon atoms are omitted for clarity except for the imino fragment.

coordination of diphenylsilane to complex **1** was studied followed by the two possibilities for the nitrile coordination (1) to the electrophilic silicon center and (2) to the ruthenium center keeping one Si-H bonded in the η^2 binding mode. The resulting DFT calculations indicated that the pathway involving the coordination to the silicon was favored in energy

when compared to the coordination to ruthenium (SI, Figure S55). The proposed mechanism is shown in Scheme 4.

This preliminary coordination of the nitrile to the Si atom is supported by the LUMO of complex **3**, which presents a lobe centered on the silicon atom (Figure 2). The nitrile coordination leads to an elongation of the Si-H bonds in

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Notes

The authors declare no competing financial interest.

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