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# Transparent Near-IR Dye-Sensitized Solar Cells: Ultrafast Spectroscopy Reveals the Effects of Driving Force and Dye Aggregation

Mate Kurucz<sup>+, [a]</sup> Ilias Nikolinakos<sup>+, [a]</sup> Jimmy Soueiti,<sup>[a]</sup> Thibaut Baron,<sup>[b]</sup> Fionnuala Grifoni,<sup>[c]</sup> Waad Naim,<sup>[c, d]</sup> Yann Pellegrin,<sup>[b]</sup> Frédéric Sauvage,<sup>\*[c]</sup> Fabrice Odobel,<sup>\*[b]</sup> and Stefan Haacke<sup>\*[a]</sup>

In the context of developing transparent near-IR absorbing dye-sensitized solar cells, diketopyrrolopyrrole (DPP) cyanine dyes have recently emerged as an alternative to strongly aggregating linear cyanines. In our efforts to increase both the power conversion efficiency (PCE) and the average visible transmittance (AVT), a thienylated version, called **TB202**, that shows a red-shifted absorption with respect to our champion dye **TB207** was designed. However, the lower energy LUMO level of **TB202** brings along a lower driving force ( $-\Delta G$ ) for carrier injection, which we recently identified as the main parameter limiting the PCE to 1.5% in the best device conditions. In the

present paper, we publish a detailed account of the effect of the de-aggregating cheno-deoxycholic acid (CDCA) for both **TB207** and **TB202**. Both transient absorption (TAS) and fluorescence up-conversion (FLUPS) data are presented, which allow to quantitatively compare the effect of  $-\Delta G$  and the CDCA concentration, in terms of the kinetic competition of ensemble averaged carrier injection and monomer-to-aggregate energy transfer (ET) rates. A comprehensive picture emerges on how ET is reduced by higher CDCA concentrations, leading in the best device conditions to injection efficiencies in the range of 65% for **TB207** and only 35% for **TB202**.

## Introduction

It goes without saying that photo-voltaics is a key technology for renewable and sustainable production of electricity and solar fuels. R&D in this area is extremely active with the aim of combining highest power conversion efficiencies (PCE), with low production costs and long-term durability of the solar cells

(SC). Third generation solar cells, based on perovskites, organic molecular systems or hybrid solutions, such as dye-sensitized solar cells (DSSCs) need to mature more in terms of long-term stability, but they are already explored for alternative modes of usage (e.g. integration in tissue) or large scale deployment approaches, like “building-integrated PV”.<sup>[1,2]</sup> For the latter in particular, a very appealing prospect is the development of transparent and colorless solar cells, the active materials of which absorb only the invisible near-UV<sup>[3]</sup> and near-IR parts of the sun’s spectrum. Since 45% of sunlight radiation received on Earth lies in the near-IR ( $\lambda \leq 1700$  nm), this part of the spectrum is presently targeted by the development of selectively near-IR absorbing solar cells (SCs), based either on organic photo-voltaics (OPV)<sup>[1,4]</sup> or on DSSCs.<sup>[5–7]</sup> The transparency of the SCs is evaluated by the “average visible transmission” (AVT). For aesthetic reasons, the transparent see-through SCs should not alter the colors of transmitted images, which is quantitatively evaluated by the “color rendering index” (CRI). The most successful devices combine all three metrics, or maximizes at least the product  $AVT \times PCE = LUE$ , the “Light Utilization Efficiency”.<sup>[8]</sup> The state-of-the-art in the development of materials (dyes, electrolytes, anodes, etc.) for transparent DSSCs was recently and comprehensively reviewed.<sup>[9]</sup>

Increasing AVT goes at the expense of PCE, in particular for non-wavelength selective absorbers: The theoretical PCE limit of a 100% AVT single junction cell using wavelength-selective transparent photovoltaics (TPV) technology, i.e. combining UV-selective and NIR-selective absorption, is 20.5%,<sup>[8]</sup> lower than the Schottky-Queisser limit for UV/VIS/near-IR absorbing semiconductors. In practice, a lower 10.8% can only be expected after considering all optical losses from the selective conversion

[a] Dr. M. Kurucz,<sup>+</sup> Dr. I. Nikolinakos,<sup>+</sup> J. Soueiti, Prof. S. Haacke  
Institut de Physique et Chimie des Matériaux (IPCMS)  
University of Strasbourg – CNRS  
23, Rue du Loess  
67034 Strasbourg, France  
E-mail: stefan.haacke@ipcms.unistra.fr

[b] Dr. T. Baron, Dr. Y. Pellegrin, Dr. F. Odobel  
Nantes Université – CNRS, CEISAM, UMR 6230  
44000 Nantes, France  
E-mail: Fabrice.Odobel@univ-nantes.fr

[c] Dr. F. Grifoni, Dr. W. Naim, Dr. F. Sauvage  
Laboratoire de Réactivité et Chimie des Solides  
Hub de l'énergie,  
Université de Picardie Jules Verne – CNRS  
15 Rue Baudelocque  
80000 Amiens, France  
E-mail: frederic.sauvage@u-picardie.fr

[d] Dr. W. Naim  
Present affiliation: Dept. Chem. Eng. & Mat. Sc., Michigan State University,  
East Lansing, USA

[†] These authors contributed equally.

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of the UV region and a reasonable long-wavelength limit of 950–1000 nm for NIR dyes.<sup>[8]</sup>

Among the wavelength-selective technologies, organic solar cells afford the highest LUE values of 2.2–3.7%, with PCEs ranging between 4% and 10%, and AVTs between 40 and 60%.<sup>[1,4]</sup> Similar figures were achieved with DSSCs based on a squaraine dye (coded HSQ5)<sup>[10]</sup> with a 3.66% PCE and a max. transmittance of 60% at 560 nm. When the absorbance is further shifted into the near-IR part with cyanine dyes, we recently demonstrated a combination of PCE 3.1% and AVT 76% without any visible coloration<sup>[5]</sup> and AVT reaching above 80% using thioate-based electrolyte.<sup>[11]</sup> Since, these cyanines have a strong tendency to form excited state quenching aggregates,<sup>[12–16]</sup> our most recent development explored more bulky diketopyrrolopyrrole (DPP) cyanines, such as **TB207**, the chemical structure of which is displayed in Scheme 1. Here the best figures of merit were PCE=3.9% and AVT=76%.<sup>[6]</sup> With the aim of red-shifting the absorption spectrum further, and possibly increasing the electron injection driving force  $-\Delta G$ ,<sup>[17,18]</sup> **TB202** was synthesized with thienyl groups at the place of phenyls at the bipyrrrole unit (Scheme 1). **TB202** displays indeed a 30 nm-red shift of the absorption spectrum, and an improved AVT, but **TB202**'s LUMO level is only slightly above the TiO<sub>2</sub> conduction band edge, resulting in a small driving force of  $-\Delta G=0.02$  eV for electron injection,<sup>[7]</sup> as compared to 0.10 eV for **TB207**.<sup>[6]</sup> These values determined by electrochemistry (cyclic voltammetry) for the dyes in solution do not represent the real values in the TiO<sub>2</sub> devices.<sup>[19]</sup> Indeed, the electronic energy levels of the TiO<sub>2</sub> conduction band down-shift due to the presence of 1 M Li<sup>+</sup> in the electrolyte by more than 100 meV,<sup>[20–22]</sup> so that the nominal  $\Delta G$ 's become more negative for both dyes. But, since this down-shift does not depend on the nature of the dye adsorbed on the TiO<sub>2</sub>, we assume that it is the same for both **TB207** and **TB202**.

In addition, and as it is well documented in the literature, dyes with a strong dipole moment have an additional effect on the energy of the conduction band that depends on the dipole direction.<sup>[23,24]</sup> The energy is decreased when the dipole moment is pointing away from the TiO<sub>2</sub> surface (positive dipole) and conversely when it is directed in the opposite direction. However, since **TB202** and **TB207** have a symmetric core structure their dipole moment is negligible both in the ground

and excited state. A remaining but negligible small dipole remains due to the acid group grafting the dyes on TiO<sub>2</sub>. In conclusion, it is reasonable to assume that the difference of the  $\Delta G$  values determined by electrochemistry (80 meV larger for **TB207**) is valid for device conditions as well.

This is reflected by the fact that the short-circuit current  $J_{sc}$  and the PCE are 2.5-times lower for **TB202** than for **TB207**, in the best device conditions.<sup>[7]</sup> Femtosecond transient absorption and spectroscopy showed that the carrier injection rate (monomer-to-TiO<sub>2</sub>, CT) in **TB202** is indeed significantly smaller than in **TB207**, and that the competition with monomer-to-aggregate energy transfer (ET) limits the injection efficiency to  $\leq 40\%$  on the sub-ns time scale.<sup>[7]</sup>

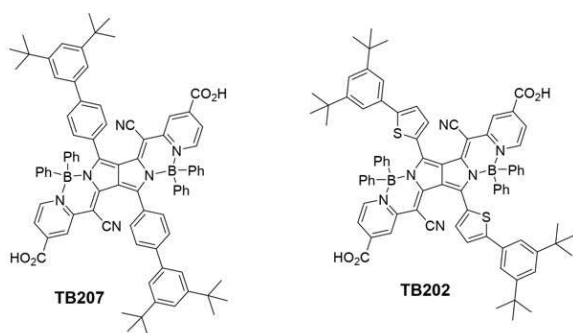
Ultrafast UV/VIS/near-IR spectroscopy is the method of choice for the study of the photo-induced processes in DSSCs as it provides a direct insight into electron injection rates and efficiencies, as well as possible geminate recombination processes.<sup>[25]</sup> The reader is referred to sound review articles<sup>[15,25–27]</sup> as well as to the early papers reporting on ultrafast spectroscopy in spectral domains ranging from UV/VIS to THz, and on a variety of organic and organometallic dyes.<sup>[28–34]</sup> Of particular interest is the detrimental effect of dye aggregates on the injection rates and yields, which was already reported for red- or near-IR absorbing dyes.<sup>[12–14,35–38]</sup>

In the present paper, we present a detailed comparison of the ultrafast spectroscopy<sup>[15]</sup> of both dyes isolated as monomers in solution, or as a heterogeneous monomer-aggregate mixture in full devices grafted on the TiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub> photo-electrodes. For the latter the spectroscopic and kinetic properties of the aggregates are of particular interest, but they were less scrutinized before. Hence, we present a detailed comparison of monomer-to-aggregate energy transfer kinetics and its competition with electron injection by monomers as a function of the concentration of the de-aggregating cheno-deoxycholic acid (CDCA). Femtosecond transient absorption (TAS) demonstrates indeed a crossover from ET to CT-dominated photo-physics as the CDCA concentration is increased. A particular point of interest is to understand how excited aggregates decay. Their spectral properties revealed by transient absorption suggest an inter-dimer charge transfer state to exist, as observed for other pyrrolopyrrole systems.<sup>[38]</sup> Finally, we discuss the possibility of electron injection, i.e. photo-current production by excited aggregates/dimers.

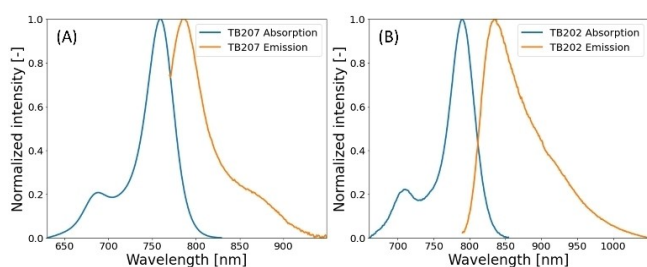
## Results and Discussion

### Photo-physical properties of dye solutions

Figure 1 shows the steady-state absorption and fluorescence spectra of **TB202** and **TB207** in EtOH. As anticipated in the introduction, the thienyl units shift the absorption maximum by 30 nm into the red. The molar extinction coefficient is high ( $> 1.3 \times 10^5 \text{ M}^{-1} \text{ cm}^{-1}$ ), due to  $\pi-\pi^*$  transitions,<sup>[6]</sup> and almost the same for both dyes. The fluorescence spectra are Stokes-shifted by 45 nm for **TB202** and 28 nm for **TB207**. **TB207** displays a good mirror-image symmetry with respect to the absorption



Scheme 1. Chemical structures of the DPP cyanine dyes **TB207** (left) and **TB202** (right).



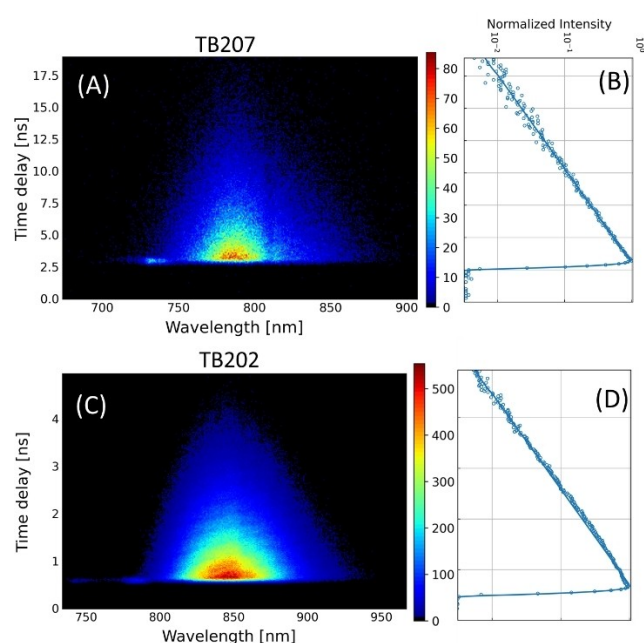
**Figure 1.** Normalized absorption and emission spectra of **TB207** (A) and **TB202** (B) in EtOH. The wavelength of maximum absorption is 30 nm red-shifted for **TB202**.

spectra, as expected for the rigid cyanide pyrrolopyrrole core of these dyes, on which HOMO and LUMO levels are localized.<sup>[6,7]</sup> The deviation in the vibrational progressions of absorption and emission is larger for **TB202**, indicating a more flexible conjugated structure.

The fluorescence decay times were measured with a time-gated streak camera (Hamamatsu C10627), as displayed in Figure 2A and B, for **TB207** and **TB202**, respectively.

We obtain the excited state lifetimes (ESL) by fitting a multi-exponential decay function convolved by the impulse response function of the system:

$$I(t) = \left[ \sum_{k=0}^n A_k \exp\left(-\frac{t}{\tau_k}\right) \right] * IRF(t) \quad (1)$$



**Figure 2.** Time-gated streak camera images of **TB207**(A) and **TB202** dyes (B) in EtOH. Fluorescence detected in magic angle configuration. Note the difference in time scales for **TB207** (20 ns, 0.2 ns resolution) and **TB202** (5 ns, 50 ps resolution). The normalized kinetic traces at the emission maximum are plotted for **TB207** (B) and **TB202** (D) with open circles for the data points and a straight line for the fitted mono-exponential function.

The IRF is the impulse response function of the setup, approximated by a Gaussian function.

The decay kinetics are wavelength-independent and are best fitted with a single time component. Based on this fitting the fluorescence lifetime is  $3.1 \pm 0.015$  ns for **TB207** dye solution and  $0.82 \pm 0.01$  ns for **TB202**. In contrast to **TB207**, **TB202**'s excited state lifetime is sub-ns, which is most likely due to internal conversion, in line with a more flexible structure. Indeed, as transient absorption data show (see below), no photo-product such as a triplet state is observed for ns delays.

Additional experiments were performed for both dyes in EtOH solution with broadband fluorescence up-conversion (150 fs time resolution), which highlight a progressive red-shift of the fluorescence spectrum due to excited state relaxation on a  $\approx 30$  ps time scale (see Figure S1 and S10). As an example, for **TB202** (Figure S10), a two-component exponential fitting was performed, which yielded lifetimes of 37 ps and 570 ps. The long lived component is similar to the value obtained by Streak camera (Table 1).

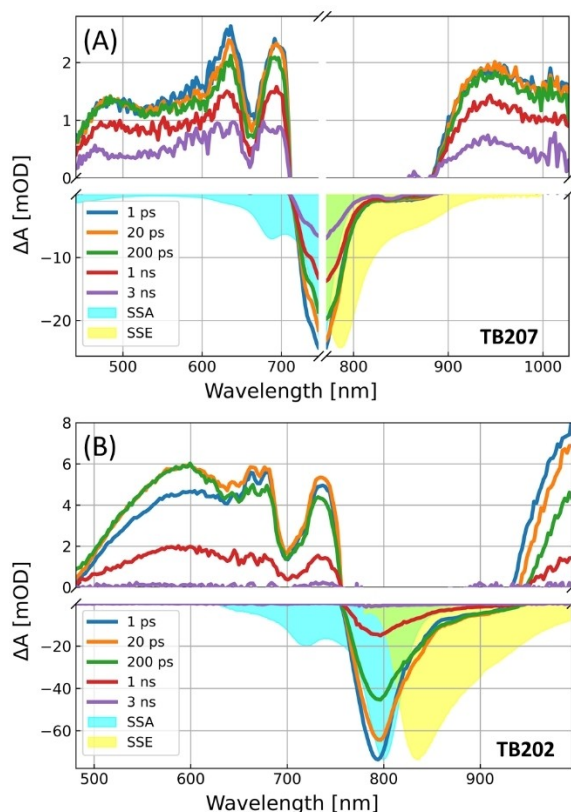
Most importantly, these experiments show that the intrinsic exc. state lifetimes (ESL) of these new dyes are much longer than usual lifetimes for charge injection ( $\leq 50$  ps),<sup>[6]</sup> so that the spontaneous radiative or non-radiative  $S_1-S_0$  relaxation is not in competition with the latter process central for the DSSC operation. In addition to fluorescence decay times, it is important to record reference  $\Delta A$  spectra for the isolated dyes in order to identify the characteristic bands and transition energies pertaining to  $S_1-S_n$  excited state absorption. Figure 3 presents the time-resolved TAS spectra at selected delay times for **TB207** (a) and **TB202** (b), obtained for excitation with 50 fs pulses at 750 and 770 nm, respectively. A wavelength interval of  $\pm 5$  nm around the pump peak wavelength is not represented for **TB207** due to residual pump light scattering. The shape of the  $\Delta A$  spectra does not depend on the delay time, only their amplitude does in agreement with the ESL determined by picosecond fluorescence (see above).

One can notice that the positive  $\Delta A$  due to excited state absorption (ESA) is typically 10 times smaller in amplitude than the negative  $\Delta A$ , due to ground state bleach (GSB, 710–800 for **TB207** and 760–820 nm for **TB202**) and stimulated emission (SE,  $> 760$  nm for **TB207** and  $> 800$  nm for **TB202**). The ESA spectrum most likely extends over the whole wavelength range

**Table 1.** Fluorescence decay times of **TB202** and **TB207** in EtOH measured by Streak camera and **TB207** cells measured with fluorescence up-conversion. For the data from  $\text{TiO}_2$  and  $\text{Al}_2\text{O}_3$  cells, the lifetimes were obtained by fitting a modified stretched exponential function (eq. 2).<sup>[39]</sup> A typical error of 3% applies to all quantities.

Sample	$\tau_0$	$\beta$	Av. decay rate $k$ (ns <sup>-1</sup> )
<b>TB207</b> Solution (Streak camera)	3.1 ns	–	0.323
<b>TB202</b> Solution (Streak camera)	0.82 ns	–	1.22
<b>TB207</b> $\text{Al}_2\text{O}_3$ 0 mM CDCA (Up-conversion)	3.8 ps	0.3	79
<b>TB207</b> $\text{TiO}_2$ 5 mM CDCA (Up-conversion)	2.9 ps	0.4	138





**Figure 3.** TAS of TB207 (A) and TB202 (B) dyes in EtOH. TAS spectra are compared to the inverted spectrum of the steady-state absorption (SSA) and steady-state emission (SSE). A positive sign signal indicates increased excitation-induced absorbance, while negative signals are due to either decreased absorbance or stimulated emission.

(440–1030 nm). The comparison of the SE amplitude with the steady-state fluorescence spectrum (SSE) shows that the former is reduced in the range of 800–900 nm (TB207) or 850–950 nm (TB202), indicating a positive ESA at these wavelengths. It is likely, but not certain that a similar compensation acts also in the GSB region. We note that ESA shows up in these spectra as two positive bands extending from 440 to 700 and from 890 to 1030 nm, and we expect these same features and band shapes to be present in the DSSC TAS data.

The ESL lifetimes obtained for TB207 from the fits ( $34 \pm 3$  ps and  $2.7 \pm 0.5$  ns) of these data are in good agreement with the values obtained from time-resolved fluorescence (table 1). The shorter 34 ps component is due to excited state relaxation, also observed by fluorescence up-conversion with 150 fs time resolution (see SI, discussed above), and the second one has a large error bar due to the limited time range scanned (4 ns).

### Aggregation of dyes grafted on TiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub>

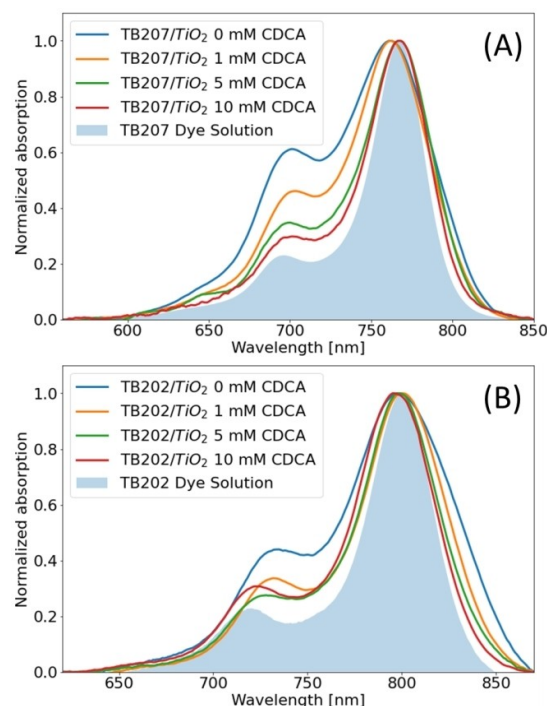
Proto-type DSSCs devices with  $10 \times 10$  mm<sup>2</sup> active areas were assembled with electrolyte composition and other material details described in the “experimental section”. A 0.1 mM dye solution was used for the grafting of both the TB202 and TB207 samples. Because both dyes are perfectly symmetrical,

they do not exert any noticeable interfacial dipole moment which could influence the electrostatic equilibrium with the TiO<sub>2</sub> density of states. This is also verified experimentally by charge extraction technique showing that neither the distribution of density of states nor the energetics of the TiO<sub>2</sub> conduction band edge is affected by the structural change of two dyes (Figure S14).

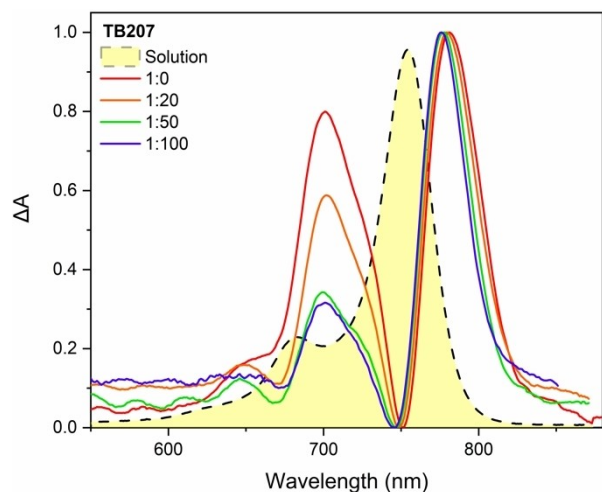
Aggregation of dyes leads to broadening of the absorption spectra as highlighted in Figure 4 for TB202 and TB207 on TiO<sub>2</sub>. The spectra of the Al<sub>2</sub>O<sub>3</sub> cells are given in Figures S2. Co-adsorbing cheno-deoxycholic acid (CDCA) increases the average distance between monomers and thus reduces the aggregate concentration as indicated by a reduced absorption on the high and low energy sides for higher CDCA concentration and in comparison with the aggregate-free solution spectra. In the following, we will refer to the aggregates as dimers since they are statistically the most abundant form (monomer absorption is the major component in the spectra, even for 0 mM CDCA, Figure 4).

We note that these dimer-related features are more pronounced for grafting on TiO<sub>2</sub> than for Al<sub>2</sub>O<sub>3</sub> for TB207 (see Figure S2), which is an important point to keep in mind when cells with equal aggregate concentrations are to be compared (*vide infra*). Figure 5 displays the aggregate-only absorption bands obtained from total spectra after subtraction of the one of solution, scaled with the arbitrary assumption that the difference should be zero at the peak of the monomer absorption.

It is important to underline that both dimer transitions are optically active, which suggests that the monomers are most likely in an oblique side-by-side arrangement with a certain tilt



**Figure 4.** Normalized absorption spectra of TB207(A) and TB202(B) TiO<sub>2</sub> solar cells. Blue areas are the steady-state absorption spectra of the solution.



**Figure 5.** Aggregate absorption spectra ( $\Delta A = A_{\text{DSSC}} - A_{\text{solution}}$ ) extracted from the total TB207/TiO<sub>2</sub> DSSCs, for increasing CDCA concentration by subtracting the corresponding absorption spectra in solution. Dashed line is the steady-state monomer absorption spectrum (solution).

angle  $\alpha$  between them.<sup>[40,41]</sup> This is the most reasonable scenario, since the anchoring groups of all dyes are located on the molecular edges, implying a grafting configuration in which the dyes are exposed to interaction with the parallel planes of the neighboring ones (excluding J-aggregate arrangement,  $\theta = 0^\circ$ , see SI). We note that when these difference spectra are normalized to the maximum of the low energy aggregate peak (Figure 5), for higher CDCA concentrations the high energy aggregate band loses intensity. This indicates that  $\alpha$  is changing as a function of CDCA concentration.

A quantitative analysis of the aggregate spectra based on Kasha's theory<sup>[18,42,43]</sup> shows for TB207 that  $\alpha$  increases from  $85^\circ$  at zero CDCA concentration to  $\approx 110^\circ$  for 10 mM of CDCA (see SI). This is most likely due to an increasing distance between the COO<sup>-</sup> anchoring groups on the surface of the SC nanoparticles since for higher concentrations more CDCA intercalates. The average monomer distance in a dimer, measured between the molecular centers, is CDCA concentration-independent: 18.5–19 Å, but this is in agreement with the observation that the energy splitting, the Davydov splitting, expressing the distance-dependent dipole-dipole is constant for all CDCA concentrations. Finally, we note that the low energy absorption band of the aggregates with a peak at 790 nm is perfectly in resonance with the emission spectrum of the monomers, thus optimizing the rates for monomer-to-aggregate energy transfer (ET).

### Effect of aggregation on excited state dynamics of TB207 grafted on TiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub>

The ET process is evidenced by a reduction of the fluorescence lifetime when TB207 is grafted on Al<sub>2</sub>O<sub>3</sub>. Since the conduction band edge of Al<sub>2</sub>O<sub>3</sub> is  $\approx 0.41$  eV higher in energy than for TiO<sub>2</sub>, electron injection is impossible (the LUMO of TB207 is only

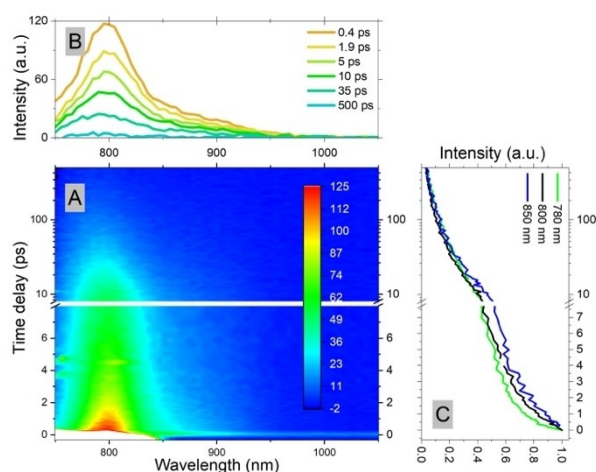
$-\Delta G = 0.1$  eV above the CB edge of TiO<sub>2</sub>) and the reduced ESL is due to ET only. The latter effect considerably reduces the ESL, as shown in Figure 6 for TB207/Al<sub>2</sub>O<sub>3</sub> with 0 mM CDCA, where the fluorescence intensity is reduced by 50% within 5 ps. The time-resolved spectra (panel B) are almost delay-independent, apart from a slight 5–10 nm redshift in the first 5 ps, due to excited state relaxation. The kinetic traces in panel C show consistently a faster initial decay for the 780 and 800 nm traces. Due to the structural heterogeneity of the monomer-to-aggregate distances and relative orientations, ET rates are expected to span a large range of values. We approximate these decays with a modified version of the standard stretched exponential function<sup>[39]</sup> described by equation (2):

$$I(t) = \exp\left(1 - \left(1 + \frac{t}{\tau_0}\right)^\beta\right) * IRF(t) \quad (2)$$

Eq. (2) corrects a major failure of the classical Kohlrausch function, for which the average decay rate  $\langle k \rangle = k(t=0)$  diverges.<sup>[39]</sup> With eq. (2), it follows  $\langle k \rangle = k(t=0) = \frac{\beta}{\tau_0}$ .<sup>[39]</sup> Furthermore, the fitting parameter  $\tau_0$  is related to the ensemble average ESL by  $\langle \tau \rangle = \tau_0 \frac{e}{\beta} \Gamma\left(\frac{1}{\beta}, 1\right)$ <sup>[39]</sup> where  $\Gamma$  is the incomplete gamma function.

Table 1 gives the fitting parameters in detail, for the 800 nm trace, and also for the 5 mM CDCA/TiO<sub>2</sub> device. We note that due to monomer-to-dimer ET in the TB207/Al<sub>2</sub>O<sub>3</sub> device,  $k$  is increased by more than a factor of 100, as compared to the TB207 solution. For comparison, Table ST1 gives the result of a 3-exp. fit (average lifetime  $\langle \tau \rangle = 54$  ps).

Femtosecond TAS reveals then the spectral signature of ET, namely the absorption spectrum of the excited dimers (Agg\*).



**Figure 6.** Broadband fluorescence up-conversion data for TB207/Al<sub>2</sub>O<sub>3</sub> with 0 mM CDCA, and excited at 730 nm. A: Time- and wavelength-resolved data; B: Emission spectra as a function of delay; C: Normalized fluorescence kinetics at selected wavelengths. Data are averaged within  $\pm 5$  nm around the central wavelength indicated. Note the break in the time axis at 8 ps. The scale is logarithmic thereafter. The Al<sub>2</sub>O<sub>3</sub> layer induces very intense laser pump pulse scattering which shows up as the white area ( $> 1000$  cts) for delays up to 0.3 ps, masking the initial fluorescence. Internal reflections of the pump pulse show up as weaker replica at 3.7 and 4.6 ps.

To this end, Figure 7A shows the time-resolved  $\Delta A$  spectra for TB207/0 mM-CDCA/ $\text{Al}_2\text{O}_3$  (2 and 5 mM in Figure S3).

At short delay times (< 1 ps) the positive  $\Delta A$  at < 700 and > 850 nm can be attributed dominantly to monomer ESA, due to its resemblance with the corresponding spectra obtained for the dyes in solution (Figure 3). Fluorescence data indicated that the excited state decays by  $\approx 60\%$  within 10 ps (Figure 6). The near-IR part of ESA(mono) and the GSB decay indeed within that time interval, and a prominent absorption band rises, with a peak at 870–880 nm, which we assign to ESA from aggregates, ESA(Agg\*). While the spectrum in the VIS hardly changes in amplitude it slightly blue-shifts in the first 10 ps, i.e. during the rise time of ESA(Agg\*). The remaining ESA with a peak at 620 nm then decays like the ESA at 880 nm, meaning that both features characterize ESA(Agg\*). The decay of Agg\* occurs on a few hundreds of ps. Kinetic traces at selected wavelengths are plotted in Figure S4 for TB207/ $\text{Al}_2\text{O}_3$  at different CDCA concentrations, together with the FLUPS decay kinetics for 0 and 5 mM CDCA. Since fitting with stretched exponentials is not adequate for TAS data with overlaying species- or state-specific bands, we simply perform multi-exponential fits for selected wavelengths. The decay times are summarized in Table S1, indicating a difference of the SE/GSB trace with the fluorescence decay times for 0 and 5 mM CDCA. This is due to the former probing partly the slower ground state recovery since the Agg\* lifetime is longer than the monomer fluorescence. Clearly, for higher CDCA concentrations the ET transfer process slows down and the amplitude of ESA(Agg\*) decreases due to the lower concentration of aggregates (see Figure S3 for the spectra with 2 and 5 mM CDCA). When we assume that all the spectral features decay with the same set of four time constants (global fit), we obtain the decay-associated difference spectra (DADS), which are presented for 0 mM CDCA in Figure S5A together with a schematic allowing to identify the spectral signatures from monomer and aggregate excited states.

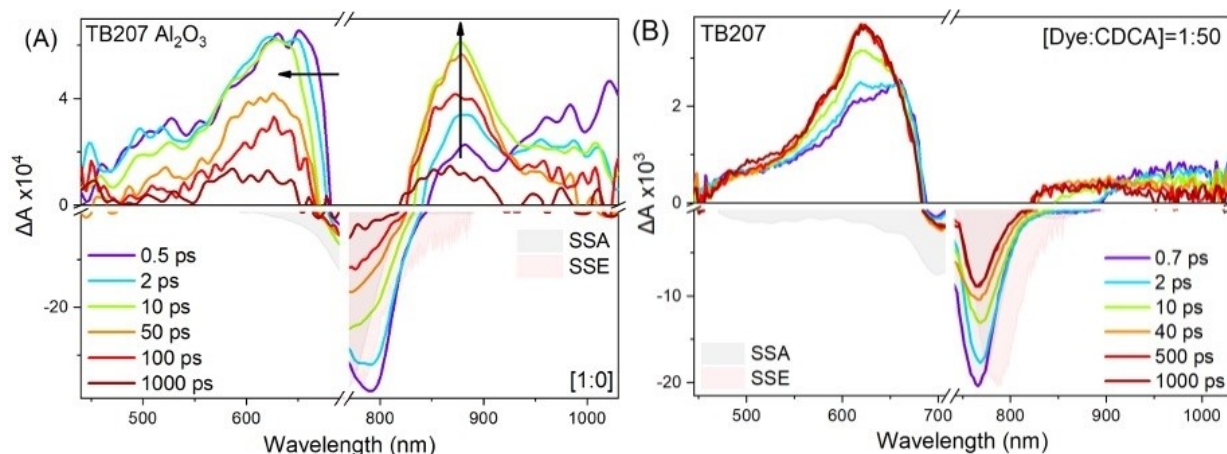
From these global fits, as well as from the fits of individual kinetic traces (table ST1), it is clear that ET transfer is slowed down for higher CDCA concentrations since the average monomer-to-dimer distance increases. The excited dimers then decay with the lifetimes  $\tau_3$  and  $\tau_4$  (Table ST1 and global fit Figure S5), which are in 100 ps – 2 ns range. Molecular dimers are known to form intra-dimer CT states,<sup>[16,44–46]</sup> and the resemblance of the red part (550–650 nm) of the ESA(Agg\*) with the spectrum of the TB207 + cation (see below) points to it. The most likely explanation is that  $\tau_3$  and  $\tau_4$  would simply be the CT recombination times.

For the TB207/ $\text{TiO}_2$  devices, the central question is then on which time scale carrier injection occurs, and if it is in kinetic competition with the average rate of ET.<sup>[12,13,15]</sup> By inspection of the CDCA-dependent absorption spectra (Figure 4), we see that the aggregate concentrations differ for both  $\text{TiO}_2$  and  $\text{Al}_2\text{O}_3$  at a given CDCA concentration.<sup>[6]</sup> Indeed,  $\text{TiO}_2$  seems to have a higher binding affinity for TB207, which leads to a higher propensity of aggregate formation. We will focus now on 5 mM CDCA on  $\text{TiO}_2$ , which was identified as optimal for achieving the highest PCE values.<sup>[6]</sup> On  $\text{Al}_2\text{O}_3$  an equivalent aggregate-to-monomer ratio is found for 0 mM CDCA.<sup>[6]</sup>

The fluorescence kinetics of both devices were analyzed using the modified Kohlrausch decay (cf. eq.(2)). The injection efficiency is given by the ratio of the ensemble-averaged rates for ET vs carrier injection (CT):

$$\eta_{inj} = 1 - \eta_{ET} = \frac{\langle k_{ET} \rangle}{\langle k_{inj} \rangle + \langle k_{ET} \rangle}. \quad (3)$$

Assuming the same  $\langle k_{ET} \rangle$  for the 5 mM-CDCA/ $\text{TiO}_2$  and the 0 mM-CDCA/ $\text{Al}_2\text{O}_3$  devices,  $\langle k_{inj} \rangle + \langle k_{ET} \rangle$  is determined from the fluorescence decay of the former, and  $\langle k_{ET} \rangle$  from the latter (Table 1). With this approach, we find  $\eta_{inj} = 0.64 \pm 0.05$ , in good agreement with the value obtained previously based on multi-exponential fits.<sup>[6]</sup> However, a different approach based on *temporally integrating* the time-dependent rates yields



**Figure 7.** Time-resolved transient absorption spectra of TB207 with 0 mM CDCA grafted on  $\text{Al}_2\text{O}_3$  (A) and grafted on  $\text{TiO}_2$  with 5 mM CDCA. Excitation wavelengths 730 nm. Note the break on the x-axis at a wavelength interval of excessive pump light scattering, and the different scales on the y-axis for the negative and positive  $\Delta A$ . The shaded SSA and SSE spectra are the sign-inverted steady-state absorption and fluorescence spectra. Excitation with average density of < 40  $\mu\text{J}/\text{cm}^2$ .



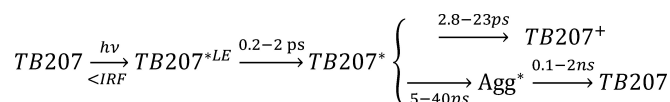
a lower injection efficiency of only  $0.44 \pm 0.05$ , as detailed in the Supp. Info. As outlined in Ref. [39] time-averaged (using the fluorescence decay as a distribution function) and ensemble-averaged rates are not the same. The relevant quantities for our discussion in the presence of structural heterogeneity are the ensemble-averaged rates, meaning that  $\eta_{inj} = 0.64 \pm 0.05$ , is the correct value.

In Figure 7B, the TAS spectra of **TB207/5 mM-CDCA/TiO<sub>2</sub>** show indeed a very prominent signature/absorption of the dye cation (TB207<sup>+</sup>), with a peak at 625 nm. Note that its amplitude  $\Delta A$  is completely formed at 40 ps and constant thereafter. In the near-IR (>850 nm), a weak and broad absorption is observed with the same kinetics, namely an increase of this absorption, or a transition from monomer-ESA to cation absorption within  $\approx 10$  ps, and a constant signal amplitude for delays  $\geq 40$  ps, indicating absence of charge recombination on the timescale probed (3.5 ns). Analysing the ratio of GSB at time-zero and long-time delays at 770 nm ( $\frac{\Delta A(t > 3ns)}{\Delta A(t=0)}$ ), one can infer an injection yield of  $\eta_{inj} = 50 \pm 10\%$ , lower than the value obtained above (fluorescence decay).<sup>[6]</sup> However, the determination by TAS is more inaccurate since the long-delay GSB amplitude is overlaid by cation absorption, and the time-zero GSB by stimulated emission, the amount of which can only be estimated.

Unlike the case of **TB207/0 mM-CDCA/Al<sub>2</sub>O<sub>3</sub>**, the Agg\* signatures do not show up, apart from a weak absorption at 850–900 nm in the 10 ps spectrum. This is line with the fact that CT wins over ET (see above), but the small amplitude of ESA(Agg\*) is surprising.

For lower CDCA concentrations (0 and 2 mM), the TAS spectra and kinetic traces can be found Figure S6 in the SI. For 0 mM CDCA, the ESA(Agg\*) appears within the first ps, already present at a delay of 0.7 ps, attesting for the ET transfer process to dominate carrier injection. Nevertheless, the cation-type temporally constant absorption spectra are observed for  $t \geq 200$  ps, when the ESA(Agg\*) contribution has decayed. Figure S7 shows the CDCA-dependent kinetic traces for **TB207/TiO<sub>2</sub>**, for characteristic wavelengths, such as 620 nm (cation max.), 780 nm (GSB/SE), 880 nm (max. of ESA(Agg\*)/tail of cation absorption). For increasing CDCA concentration, the qualitative trends of ESA(Agg\*) suppression and reduction of  $\langle k_{ET} \rangle$ , i.e. an increasing rise time of ESA(Agg\*) are apparent. We analyzed the kinetics by single wavelength and global fits, the results of which being summarized in (table SI). The decay-associated difference spectra (DADS) allow to identify the processes, which dominate for the different time scales, and as a function of CDCA. Since the excited state or photo-induced absorption spectra overlay in wavelength and time, this discussion is relatively involved and the reader is referred to the Supp. Info. for the details.

Nevertheless, the reaction scheme that emerges from the analysis of the DADS for **TB207/5 mM-CDCA/TiO<sub>2</sub>** is summarized in the scheme below with the relevant time scales (Scheme 2).



**Scheme 2.** General reaction scheme for TB207/TiO<sub>2</sub> with lifetimes obtained for 5 mM CDCA, i.e. the best device conditions. After excitation with 50 fs pulses, the local excited state TB207\*LE undergoes vibrational relaxation and forms TB207\*. Electron injection forming the TB207<sup>+</sup> cation and energy transfer yielding Agg\* compete since they occur on similar timescales (numbers taken from the global fits, see Figure S5). The latter decay back to the aggregate ground state (Agg), most probably due to recombination of the inter-aggregate CT state. TB207<sup>+</sup> cations are stable on the time scales probed ( $\leq 5$  ns). Note that the same CT/ET branching reaction scheme holds for TB202.

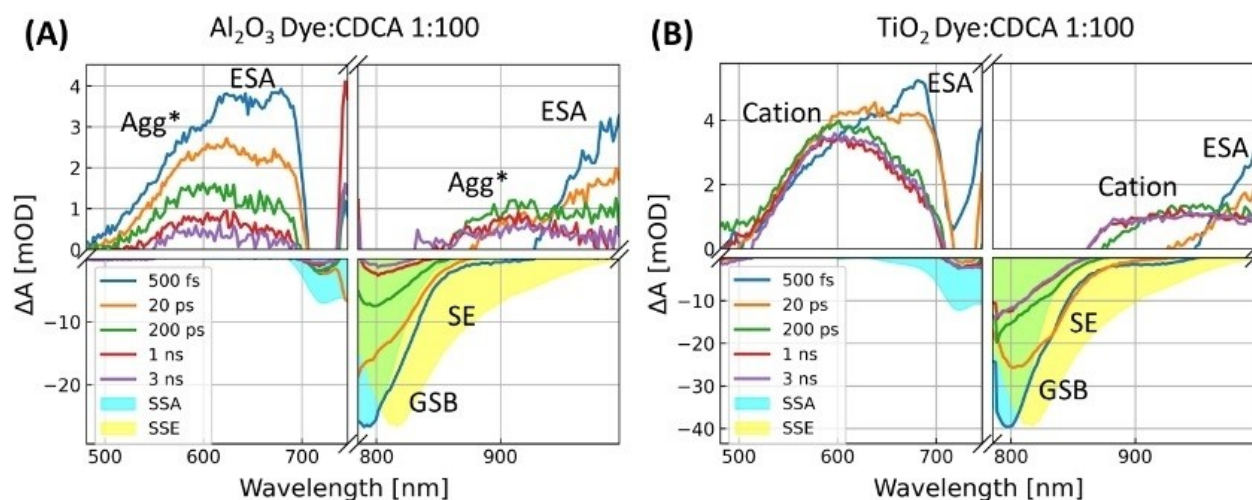
### Excited state dynamics of TB202 grafted on TiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub>: Effect of aggregation and charge transfer driving force

The femtosecond TAS was performed with higher wavelength excitation at 770 nm, due to the red-shifted absorption spectra of **TB202** compared to **TB207**. The time-resolved  $\Delta A$  spectra for **TB202/10 mM-CDCA/Al<sub>2</sub>O<sub>3</sub>** and **TB202/10 mM-CDCA/TiO<sub>2</sub>** are displayed on Figure 8A and B (0,1 and 5 mM in SI). For **TB202**, the absorption spectra on Al<sub>2</sub>O<sub>3</sub> and TiO<sub>2</sub> (Figure 4 and S2) indicate indeed similar aggregate/dimer concentrations for the same CDCA concentrations, unlike **TB207**. Ground state bleach (GSB, 795 nm) and stimulated emission (SE, 820 nm) signatures in the TAS measurement can be identified by comparison with the steady-state absorption and fluorescence spectra, respectively (plotted with inverted sign in Figure 8). Guided by the signatures of the monomer ESA signal of the **TB202** solution (Figure 3B), excited state absorption (ESA) can be identified for both DSSCs with maxima at 680 and 1000 nm. No CT can occur in Al<sub>2</sub>O<sub>3</sub> cells, consequently we can attribute the longer lifetime component (> 20 ps) to the absorption spectrum of the excited aggregate (Agg\*). The spectra do not fully decay on the time-scale of the measurement, which could indicate that ET leads to an intra-aggregate CT state,<sup>[38]</sup> which decays on a nanosecond time scale, as for **TB207**. This is supported by the spectral resemblance of  $\Delta A(Agg^*)$  and the cation differential spectra observed for **TB202/TiO<sub>2</sub>** (Figure 8B).

In the TiO<sub>2</sub> DSSC, and for delays  $\leq 20$  ps, GSB, SE and the monomer ESA occur at the same wavelengths as for Al<sub>2</sub>O<sub>3</sub> (Figure 8B). For later delays, the prominent absorption spectrum of the cation (TB202<sup>+</sup>) appears at high (10 mM) CDCA concentration as an absorption band between 520–700 nm, with a maximum at 580 nm, and a weaker flat band ranging from 870 nm into the near IR. As the dye-regeneration operates on the ms timescale only,<sup>[6]</sup> the cation spectra persist at longer timescale (> 200 ps), while the absorption spectrum of the excited aggregate decays as witnessed for the Al<sub>2</sub>O<sub>3</sub> cells. Note the slower rise time of the cation signature for **TB202**, as compared to **TB207** (Figure 7B).

Due to the spectral overlap of absorption of Agg\* and TB202<sup>+</sup>, both contributions are impossible to separate. However, the partial decay and reshaping of  $\Delta A$  in the VIS and near-IR portions in the 200 ps to 1 ns time range may indicate decay





**Figure 8.** Time-resolved transient absorption spectra of **TB202** with 10 mM CDCA grafted on  $\text{Al}_2\text{O}_3$  (A) and grafted on  $\text{TiO}_2$  with 10 mM CDCA (B). Excitation wavelengths 770 nm. A break on the x-axis was introduced at a wavelength interval of excessive pump light scattering, and different scales are used for the y-axis for the negative and positive  $\Delta A$  values. The shaded SSA and SSE spectra are the sign-inverted steady-state absorption and fluorescence spectra.

of  $\text{Agg}^*$  also for  $\text{TiO}_2$ , as observed in Figure 8A for the  $\text{Al}_2\text{O}_3$  device.

To determine the lifetimes for CT and ET at different CDCA concentrations, we performed a global analysis of the TAS data, under the assumption that the kinetics are described by a set of wavelength-independent time constants using the variable projection (VARPRO) algorithm.<sup>[47]</sup> Best agreement is obtained when the TAS of the solar cells are fitted with a minimum of 4 components and an infinite lifetime component, as shown in Table 2. The corresponding DADS are displayed in Figure S13.

These lifetimes can be associated with processes by observing the corresponding DADS (cf. Figure S13). A reaction scheme similar to **TB207** (Scheme 1) emerges, but with different lifetimes. For the  $\text{Al}_2\text{O}_3$  cells ET transfer occurs on  $\tau_2/\tau_3$  timescales. These values are increasing with the CDCA concentration, as the intermolecular distance between the monomers and the aggregates increases. The  $\text{Agg}^*$  state decays with  $\tau_4/\tau_5$  timescales back to the ground state, due to recombination of

the inter-dimer CT state. For the  $\text{TiO}_2$  cells CT and ET cannot be separated as they spectrally overlap and they both occur on  $\tau_2/\tau_3$  timescales. On  $\tau_4$  timescale  $\text{Agg}^*$  state decays, while the cation lifetime of  $\tau_5$  is longer than 3.5 ns (fitted as infinite).

Indeed, comparing these results to the lifetimes of **TB207** both ET and CT processes are slower, for comparable CDCA concentrations. The lifetimes associated with ET for  $\text{Al}_2\text{O}_3$  samples ( $\tau_2/\tau_3 = 7\text{--}77$  ps for 10 mM CDCA cell) are slower than the lifetimes of  $\text{TiO}_2$  samples when both CT and ET take place in parallel ( $\tau_2/\tau_3 = 2.8\text{--}49$  ps for 10 mM CDCA cell). The injection yield was determined for **TB202**/10 mM-CDCA/ $\text{TiO}_2$  cell to be  $\eta_{\text{inj}} = 32 \pm 10\%$ , by calculating the ratio of GSB at time-zero and long time delays at 830 nm.<sup>[7]</sup> The PCE of **TB202** is 2.5-times smaller than the PCE of **TB207**.<sup>[7]</sup> For the present 10 mM CDCA we find that ET is slower than for **TB207**, but the charge injection time is not fast enough to out-compete ET. Indeed, the main problem of the thienylated **TB202** is the lower driving force  $-\Delta G = 0.02$  eV<sup>[7]</sup> as compared to **TB207** (0.10 eV).<sup>[6]</sup>

## Conclusions

The present femtosecond experiments allow to determine the relevant time scales and average rates, related to carrier injection and monomer-to-aggregate energy transfer, for both dyes. The analysis is complicated by the structural heterogeneity of the monomer and aggregate geometries, leading to an inhomogeneous distribution of lifetimes. This is best addressed for the fluorescence decay kinetics by the modified Kohlrausch function (eq. 2), but for the TAS data a sum of exponentials has to be used. The SI show both global fits and single wavelength fits, which indicate the relevant time scales, but also the systematic errors in both approaches.

Nevertheless, the main picture that emerges is that the principal difference in PCE of devices made with both dyes comes indeed from the lower driving force  $-\Delta G$  of **TB202**, which results in carrier

**Table 2.** TAS decay times of **TB202** in  $\text{TiO}_2$  and  $\text{Al}_2\text{O}_3$  cells with different CDCA concentrations, obtained by a global fit analysis. Although the fitting uncertainties are much smaller, an 20 fs error applies for  $\tau_1$ , which is close to the time resolution. We estimate a 3–5% systematic error for  $\tau_2$  &  $\tau_3$ , and 10% for  $\tau_4$ .

Semiconductor	TB202: CDCA	$\tau_1$ (ps)	$\tau_2$ (ps)	$\tau_3$ (ps)	$\tau_4$ (ps)	$\tau_5$ (ps)
$\text{Al}_2\text{O}_3$	1:0	0.065	1.4	8.5	140	$\infty$
	1:10	0.093	2	18	280	$\infty$
	1:50	0.28	5.1	68	580	$\infty$
	1:100	0.29	7	77	570	$\infty$
$\text{TiO}_2$	1:0	0.88	5.3	48	500	$\infty$
	1:10	0.21	3.9	40	250	$\infty$
	1:50	0.21	4.4	46	410	$\infty$
	1:100	0.26	2.8	49	180	$\infty$

injection and quenching by ET to occur on the same time scale, thus severely limiting the injection efficiency to  $\approx 35\%$ . **TB202** has indeed a more red-shifted absorption spectrum affording a larger AVT, but at the cost of a reduced LUMO energy and hence a too low  $-\Delta G$ . In our devices, the effective  $-\Delta G$  is increased by the addition of 1 M Li<sup>+</sup> ions in the electrolyte, but the shift is apparently not sufficient to compensate for the intrinsic lower reduction potential of **TB202**.<sup>[7]</sup>

While in previous DSSCs the monomer-to-aggregate ET was considered to be a pure quenching, i.e. loss channel, the case of **TB207** questions this assumption. Indeed, the PCE is only weakly reduced to 3.0% at 0 mM CDCA, as compared to 3.9% at 5 mM CDCA (best device conditions), and  $J_{sc}$  is relatively large, i.e. in the range of 16 mJ/cm<sup>2</sup>, even for 0 mM CDCA.<sup>[6]</sup> However, without any CDCA, the TAS data are dominated by the ESA(Agg\*) signature and cations are emerging only at long delay times with a low  $\Delta A$  amplitude (Figure S6), indicating a significantly lower cation concentration than at higher CDCA concentrations. This indicates that the low energy dimer state might possibly have enough driving force to contribute to carrier injection, yet on a slower time scale than the monomers. Some experiments were conducted on different near-IR dyes, as a function of the excitation wavelengths, in particular exciting the low energy edge of the absorption spectra, where the aggregate concentration dominates.<sup>[48]</sup> The effect on the injection efficiency was nevertheless not convincingly detectable. Our next step will therefore be to conduct two-dimensional electronic spectroscopy with broadband excitation around 750 nm,<sup>[49]</sup> which should allow us to detect the photo-induced cation concentration precisely as a function of the excitation wavelength.<sup>[50]</sup>

## Experimental section

Steady-state absorption spectra were recorded on a LAMBDA 950 UV/Vis spectrometer from PerkinElmer. The steady-state emission was measured using an Edinburgh instrument FLS920 Fluorescence spectrometer. The streak camera model that is used in our experiments is the HAMAMATSU C10627, mounted behind a spectrograph Jobin Yvon, 25-cm focal length, 50 g/mm grating, blazed at 600 nm). The time resolution is 0.05 and 0.2 ns for streak camera delay ranges used for Figure 2. Fluorescence detected in magic angle configuration.

**Femtosecond set-ups.** For both the time-resolved transient absorption and the fluorescence up-conversion setup (shown in S15) a Ti:sapphire laser system (FemtoLasers Synergy 20) was used. The laser pulses are amplified in a Chirped Pulsed Amplification with regenerative amplifier (Pulsar-Amplitude Technologies) resulting in a 0.5 mJ pulse energy, 800 nm central wavelength, 5 kHz repetition rate pulse train. The excitation wavelength is tuned to the absorption band of the DSSC's using a TOPAS-Prime (Light Conversion). The average pulse energy density (pulse energy/ (1/e<sup>2</sup> area)) of the excitation was chosen to prevent the signatures of exciton-exciton annihilation, which was observed in devices at higher energy densities.<sup>[48]</sup> For the **TB207** solution an energy density of  $\sim 60 \mu\text{J}/\text{cm}^2$  was used, but a lower energy density of  $\sim 30 \mu\text{J}/\text{cm}^2$  was required for the **TB207** DSSCs. In case of the **TB202** dye, the exciton-exciton annihilation effect appeared to be reduced. We could use average pulse energy densities of  $260 \mu\text{J}/\text{cm}^2$  and  $\sim 80 \mu\text{J}/\text{cm}^2$  for the solution and the DSSCs, respectively. All data were recorded under magic angle conditions.

In the fluorescence up-conversion setup (see Figure. S14) the pump excites the sample and the produced fluorescence is imaged onto a BBO crystal, by a pair of parabolic mirrors. Here the sum frequency is generated with a near-IR gate pulse (signal from TOPAS-Prime). The resulting up-converted signal is then detected as a function of gate-excitation delay by a LN<sub>2</sub>-cooled CCD. The following conditions for SFG are optimized so as to obtain broadband phase-matching<sup>[51–53]</sup> ( $> 100 \text{ nm}$  for the present near-IR dyes): type II phase-matching with the gate beam in extra-ordinary polarization,  $\approx 10^\circ$  non-collinear angle between the gate beam and the center of the fluorescence cone, and a relatively thin 0.4 mm BBO crystal (cut at  $25^\circ$ ).

**Synthesis of TB207 and TB202.** The protocol for the synthesis of both pyrrolopyrrole cyanine dyes is described in detail in Refs. [6] and [7], respectively. The <sup>1</sup>H NMR and MALDI-TOF spectra are replicated in the Supp. Info of the present manuscript (Figures S16–S19).

**Sample preparation.**<sup>[48]</sup> For the steady-state studies on dye solutions, these were prepared with lower optical densities in the range of 0.1–0.3/mm, in order to avoid reabsorption in fluorescence experiments. Higher OD values are used for the TAS and FLUPS measurements (0.3–0.55/mm). Spectroscopy grade ethanol was purchased from Sigma-Aldrich and used as received.

The DSSCs were prepared for both dyes following the protocol detailed here. The anodes are TEC11-type conductive glasses coated with a fluorine doped tin oxide (FTO) layer. The  $\approx 4 \mu\text{m}$  thick semiconductor layer was deposited by screen printing. The TiO<sub>2</sub> coated anodes were heated up gradually to 500 °C in an oven, while the Al<sub>2</sub>O<sub>3</sub> covered anodes were heated up to 300 °C and kept at these temperatures for half an hour. Then they were dipped overnight (15–17 h) in a dye solution (0.1 mM dye, 90% EtOH 10% CHCl<sub>3</sub>). The cathodes are TEC11-type conductive glasses coated with FTO. Small diameter ( $< 1 \text{ mm}$ ) holes allow the electrolyte to be injected by vacuum back-filling at the final stage of assembly. Cathodes were cleaned before assembly by three consecutive ultrasonic baths (1×15 min Acetone, 2×15 min EtOH).

A transparent surlin polymer hole square (12×12 mm outer size and 10–10 mm inner size) is placed between the anode and the cathode to glue the two electrodes together on a hot plate at 120 °C melting the surlin. The resulting cavity was filled with an acetonitrile-based PV-grade electrolyte provided by G-Lyte (Amiens, France). It contains 1 M LiI (lithium iodide), 1 M DMII (1,3-dimethylimidazole iodide) and 0.3 M I<sub>2</sub> (iodine). After electrolyte injection, the holes are sealed with thin cover slip glasses, and melted surlin.

The DSSCs displayed optical densities at the peak absorbance in the range of OD = 0.4–1.0, with the lower values corresponding to 50 or 100 mM CDCA. A background OD of  $\approx 0.06$ –0.08 due to air-glass reflectivity and scattering by the semiconductor nanoparticles was subtracted.

## Supporting Information

The supporting information contain additional tables and figures, the determination of the structural parameters of the aggregate dimers, as well as the calculation of the injection efficiency for **TB207** based on time-averaged fluorescence decay rates (as opposed to the appropriate ensemble averaging). The authors have cited additional references within the Supporting Information.<sup>[54–57]</sup>

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## Conflict of Interests

The authors declare no conflict of interest.

## Data Availability Statement

The data that support the findings of this study are available from the corresponding author upon reasonable request.

**Keywords:** near-IR DSSC · pyrrolopyrrole cyanines · charge transfer · energy transfer · molecular dimers · ultrafast · transient absorption · broadband fluorescence up-conversion

- [1] R. R. Lunt, V. Bulovic, *Appl. Phys. Lett.* **2011**, *98*, 113305.
- [2] C. J. Traverse, R. Pandey, M. C. Barr, R. R. Lunt, *Nat. Energy* **2017**, *2*, 849–860.
- [3] A. D. Wulansari, D. Hayati, D. X. Long, K. Choi, J. Hong, *Sci. Rep.* **2023**, *13*, 1–11.
- [4] C.-C. Chen, L. Dou, R. Zhu, C.-H. Chung, T.-B. Song, Y. B. Zheng, S. Hawks, G. Li, P. S. Weiss, Y. Yang, *ACS Nano* **2012**, *6*, 7185–7190.
- [5] W. Naim, V. Novelli, I. Nikolinakos, N. Barbero, I. Dzeba, F. Grifoni, Y. Ren, T. Alnasser, A. Velardo, R. Borrelli, *JACS Au* **2021**, *1*, 409–426.
- [6] T. Baron, W. Naim, I. Nikolinakos, B. Andrin, Y. Pellegrin, D. Jacquemin, S. Haacke, F. Sauvage, F. Odobel, *Angew. Chem.* **2022**, *134*, e202207459.
- [7] T. Baron, W. Naim, M. Kurucz, I. Nikolinakos, B. Andrin, Y. Pellegrin, D. Jacquemin, S. Haacke, F. Sauvage, F. Odobel, *J. Mater. Chem. A* **2023**, *11*, 16767–16775.
- [8] R. R. Lunt, *Appl. Phys. Lett.* **2012**, *101*, 043902.
- [9] F. Grifoni, M. Bonomo, W. Naim, N. Barbero, T. Alnasser, I. Dzeba, M. Giordano, A. Tsaturyan, M. Urbani, T. Torres, C. Barolo, F. Sauvage, *Adv. Energy Mater.* **2021**, *11*, 2101598.
- [10] K. Zhang, C. Qin, X. Yang, A. Islam, S. Zhang, H. Chen, L. Han, *Adv. Energy Mater.* **2014**, *4*, 1301966.
- [11] W. Naim, F. Grifoni, V. Challuri, D. Mathiron, S. Ceurstemont, P. Chotard, T. Alnasser, I. Dzeba, N. Barbero, S. Pilard, C. Barolo, F. Sauvage, *Cell Rep. Phys. Sci.* **2023**, *4*, 101455.
- [12] K. Pydzińska, M. Ziólek, *Dyes Pigm.* **2015**, *122*, 272–279.
- [13] M. Ziólek, J. Karolczak, M. Zalas, Y. Hao, H. Tian, A. Douhal, *J. Phys. Chem. C* **2014**, *118*, 194–205.
- [14] V. Novelli, N. Barbero, C. Barolo, G. Viscardi, M. Sliwa, F. Sauvage, *Phys. Chem. Chem. Phys.* **2017**, *19*, 27670–27681.
- [15] C. Martin, M. Ziólek, A. Douhal, *J. Photochem. Photobiol. C* **2016**, *26*, 1–30.
- [16] H.-W. Bahng, A. Hagfeldt, J.-E. Moser, *J. Phys. Chem. C* **2018**, *23*.
- [17] R. A. Marcus, *J. Chem. Phys.* **1956**, *24*, 966–978.
- [18] Lanzani, Guglielmo, *The Photophysics behind Photovoltaics, Photonics*, Wiley-VCH Verlag GmbH, 2012.
- [19] H. Rensmo, K. Westermark, S. Södergren, O. Kohle, P. Persson, S. Lunell, H. Siegbahn, *J. Chem. Phys.* **1999**, *111*, 2744–2750.
- [20] D. F. Watson, G. J. Meyer, *Coord. Chem. Rev.* **2004**, *248*, 1391–1406.
- [21] Y. Bai, J. Zhang, Y. Wang, M. Zhang, P. Wang, *Langmuir* **2011**, *27*, 4749–4755.
- [22] J. R. Jennings, Q. Wang, *J. Phys. Chem. C* **2010**, *114*, 1715–1724.
- [23] S. Rühle, M. Greenshtein, S.-G. Chen, A. Merson, H. Pizem, C. S. Sukenik, D. Cahen, A. Zaban, *J. Phys. Chem. B* **2005**, *109*, 18907–18913.
- [24] S. S. Pandey, S. Sakaguchi, Y. Yamaguchi, S. Hayase, *Org. Electron.* **2010**, *11*, 419–426.
- [25] V. Sundström, A. Yartsev, in *Energy and Environment Series*, ed. P. Piotrowiak, Royal Society of Chemistry, Cambridge, **2013**, pp. 135–160.
- [26] C. S. Ponceca, P. Chábera, J. Uhlig, P. Persson, V. Sundström, *Chem. Rev.* **2017**, *117*, 10940–11024.
- [27] S. Ardo, G. J. Meyer, *Chem. Soc. Rev.* **2009**, *38*, 115–164.
- [28] T. Hannappel, B. Burfeindt, W. Storck, F. Willig, *J. Phys. Chem. B* **1997**, *101*, 6799–6802.
- [29] H. N. Ghosh, J. B. Asbury, T. Lian, *J. Phys. Chem. B* **1998**, *102*, 6482–6486.
- [30] G. Benkö, J. Kallioinen, J. E. I. Korppi-Tommola, A. P. Yartsev, V. Sundström, *J. Am. Chem. Soc.* **2002**, *124*, 489–493.
- [31] R. Huber, J.-E. Moser, M. Grätzel, J. Wachtveitl, *J. Phys. Chem. B* **2002**, *106*, 6494–6499.
- [32] K. Sunahara, A. Furube, R. Katoh, S. Mori, M. J. Griffith, G. G. Wallace, P. Wagner, D. L. Officer, A. J. Mozer, *J. Phys. Chem. C* **2011**, *115*, 22084–22088.
- [33] N. A. Anderson, T. Lian, *Annu. Rev. Phys. Chem.* **2005**, *56*, 491–519.
- [34] H. Némec, J. Rochford, O. Taratula, E. Galoppini, P. Kužel, T. Polívka, A. Yartsev, V. Sundström, *Phys. Rev. Lett.* **2010**, *104*, 197401.
- [35] T. Geiger, I. Schoger, D. Rentsch, A. C. Véron, F. Oswald, T. Meyer, F. Nüesch, *Int. J. Photoenergy* **2014**, *2014*, e258984.
- [36] J. Sobus, J. Karolczak, D. Komar, J. A. Anta, M. Ziólek, *Dyes Pigm.* **2015**, *113*, 692–701.
- [37] A. Kaczmarek-Kedziera, D. Kedziera, *Theor. Chem. Acc.* **2016**, *135*, 1–17.
- [38] H.-W. Bahng, A. Hagfeldt, J.-E. Moser, *J. Phys. Chem. C* **2018**, *122*, 19348–19358.
- [39] M. Berberan-Santos, E. Bodunov, B. Valeur, *Chem. Phys.* **2005**, *315*, 171–182.
- [40] A. R. Monahan, D. F. Blossley, *J. Phys. Chem.* **1970**, *74*, 4014–4021.
- [41] P. Bojarski, A. Matczuk, C. Bojarski, A. Kawski, B. Kukliński, G. Zurkowska, H. Diehl, *Chem. Phys.* **1996**, *210*, 485–499.
- [42] M. Kasha, *Radiat. Res.* **1963**, *20*, 55–70.
- [43] N. J. Hestand, F. C. Spano, *Acc. Chem. Res.* **2017**, *50*, 341–350.
- [44] Y. Hong, F. Schlosser, W. Kim, F. Würthner, D. Kim, *J. Am. Chem. Soc.* **2022**, *144*, 15539–15548.
- [45] J. Sung, A. Nowak-Król, F. Schlosser, B. Fimmel, W. Kim, D. Kim, F. Würthner, *J. Am. Chem. Soc.* **2016**, *138*, 9029–9032.
- [46] H.-W. Bahng, A. Hagfeldt, J.-E. Moser, *J. Phys. Chem. C* **2018**, *19*.
- [47] D. P. O'Leary, B. W. Rust, *Comput. Optim. Appl.* **2013**, *54*, 579–593.
- [48] Nikolinakos, Ilias, *Ultrafast Spectroscopy of Transparent Dye Sensitized Solar Cells Designed for the Near-Infrared*, PhD thesis, University of Strasbourg (2022), <https://www.theses.fr/2022STRAE005>.
- [49] M. Gueye, J. Nillon, O. Crégut, J. Léonard, *Rev. Sci. Instrum.* **2016**, *87*, 093109.
- [50] F. D. Fuller, J. P. Ogilvie, *Annu. Rev. Phys. Chem.* **2015**, *66*, 667–690.
- [51] M. Gerecke, G. Bierhance, M. Gutmann, N. P. Ernstring, A. Rosspeintner, *Rev. Sci. Instrum.* **2016**, *87*, 053115.
- [52] M. Sajadi, M. Quick, N. P. Ernstring, *Appl. Phys. Lett.* **2013**, *103*, 173514.
- [53] L. Zhao, J. Luis Pérez Lustres, V. Farztdinov, N. P. Ernstring, *Phys. Chem. Chem. Phys.* **2005**, *7*, 1716–1725.
- [54] M. Yu. Losytskyy, V. M. Yashchuk, S. S. Lukashov, S. M. Yarmoluk, *J. Fluoresc.* **2002**, *12*, 109–112.
- [55] N. R. Reddy, S. Rhodes, Y. Ma, J. Fang, *Langmuir* **2020**, *36*, 13649–13655.
- [56] C. M. Carbonaro, *J. Photochem. Photobiol. C* **2011**, *222*, 56–63.
- [57] F. del Monte, J. D. Mackenzie, D. Levy, *Langmuir* **2000**, *16*, 7377–7382.

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