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## COMMUNICATION

## A *p*-type Semi-Conducting Copper(I)-1,3-Benzenedithiolate 2D Coordination Polymer with High Seebeck Coefficient

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**A new *p*-type semi-conducting and 2D coordination polymer (CP) based on thiolate ligand and Cu(+) exhibits a record Seebeck coefficient among CPs, equivalent to that obtained for inorganic compounds. Such high value is originating from the 1D {Cu-S}<sub>n</sub> tubular inorganic network, that is connected through the bridging 1,3-benzenedithiolate ligand.**

Considering the actual global warming and our over-reliance on fossil fuels, there is a tremendous need in new energy technologies with low environmental and societal impacts. Thermoelectricity, which is the ability of a material to convert heat into electricity and vice versa, is an ideal technology that would allow transforming heat waste from industries, engines, power plants, household appliances and even human bodies into power; providing sustainable energetic cycles. To reach this goal, a thermoelectric material should have a good performance and its efficiency is evaluated by a dimensionless figure of merit,  $ZT = S^2\sigma T / \kappa$ , where  $S$ ,  $\sigma$ ,  $T$  and  $\kappa$  are the Seebeck coefficient, the electrical conductivity, the absolute temperature and the thermal conductivity, respectively. Bearing in mind the Carnot cycle, the  $ZT$  value of a material should be greater than 3 for a device efficiency.<sup>1</sup>

From the discovery of thermoelectricity phenomenon, extensive studies have been carried out on inorganic materials, mostly the chalcogenides such as Bi<sub>2</sub>Te<sub>3</sub>, Sb<sub>2</sub>Te<sub>3</sub>, PbTe and

SnSe. Those thermoelectric (TE) materials have high Seebeck coefficient, up to around 500  $\mu\text{V}\cdot\text{K}^{-1}$ , but they have high thermal conductivity, which results in a modest  $ZT$ , still below 1.<sup>2</sup> The record  $ZT$  value has been reported in 2019 for an alloy, Fe<sub>2</sub>V<sub>0.8</sub>W<sub>0.2</sub>Al, with a  $ZT$  value of 5 to 6.<sup>3</sup> In addition, these thermoelectric inorganic materials have the drawbacks of poor processability, high cost and toxicity, preventing today further sustainable applications. In the meantime,  $\pi$ -delocalized organic polymers, such as poly(3,4-ethylenedioxythiophene) polystyrene sulfonate (PEDOT:PSS) or polyaniline (PANI), have appeared as good alternatives as TE materials with good shaping opportunity, light weight, non-toxicity, low cost and more importantly low thermal conductivity ( $\sim 0.2 \text{ W}\cdot\text{m}^{-1}\cdot\text{K}^{-1}$ ).<sup>4</sup> Nevertheless, the organic materials have low Seebeck coefficient ( $\sim 30 \mu\text{V}\cdot\text{K}^{-1}$ ) resulting in still modest  $ZT$ . Thus, there is huge need to further improve the  $ZT$  value of materials to achieve the Carnot efficiency, and significant efforts have to be devoted into the development of TE materials exhibiting high Seebeck coefficient, high electrical conductivity, and low thermal conductivity. An alternative is to use hybrid materials to take advantages of the inorganic and organic networks. Thus, TE composite materials which are the mixture of two compounds, are a good strategy, but the synergy between the two entities is difficult to control and requires multistep syntheses.<sup>5</sup>

Metal Organic Frameworks (MOFs) and Coordination Polymers (CPs), constructed from metal ions bonded to organic moieties, offer a diversity of structural topologies ranging from one-dimensional (1D) to three-dimensional (3D) arrangements, with possible porosity coupled with other different functionalities, which allow fine tuning of their electronic structures, resulting in intrinsic physical properties.<sup>6</sup> Those MOFs and CPs have recently appeared as potential TE materials (Table S1).<sup>7</sup> Indeed, the 2D copper-benzenedithiolate, Cu<sub>x</sub>BHT, exhibits a record electrical conductivity up to 2000  $\text{S}\cdot\text{cm}^{-1}$ , due to the delocalized charge through the  $\pi$ -d conjugated system.<sup>8</sup> The TE properties have been evaluated for different formula of Cu<sub>x</sub>BHT with  $x = 3, 4$

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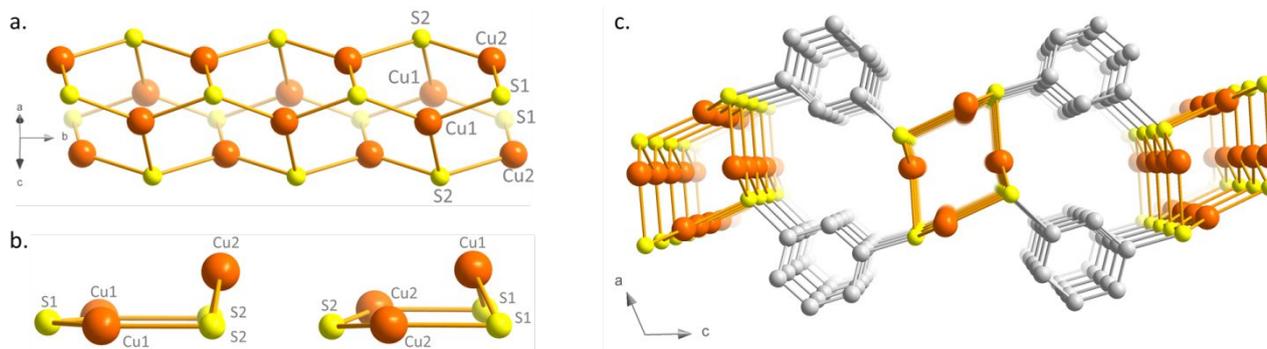
Electronic Supplementary Information (ESI) available: experimental and instrumental descriptions, SEM, TGA, FT-IR, PXRD, XPS, UV-vis and conductivity data. See DOI: 10.1039/x0xx00000x

and 5.5 and the best ZT of 0.08 at RT has been obtained with a Seebeck coefficient of  $80 \mu\text{V.K}^{-1}$  and a thermal conductivity of  $0.59 \text{ W.m}^{-1}\text{.K}^{-1}$ .<sup>9</sup> The lowest thermal conductivity of  $0.2 \text{ W.m}^{-1}\text{.K}^{-1}$  has been measured in the two  $\text{Ni}_3(2,3,6,7,10,11\text{-hexaiminotriphenylene})_2$  and  $[\text{Ni}_3(\text{perthiolated coronene})]_n$  CPs, nevertheless their low Seebeck coefficients of  $-12$  and  $47 \mu\text{V.K}^{-1}$  imply low ZT of  $1.19 \cdot 10^{-3}$  and  $0.003$ , respectively.<sup>10</sup> Zn-HAB (HAB : hexaaminobenzene), a 3D porous network, shows the best intrinsic Seebeck coefficient of  $200 \mu\text{V.K}^{-1}$  at RT and under vacuum, but this MOF presents a low conductivity and suffer of instability under  $\text{O}_2$ , due to the weak interactions between Zn(II) and the amino groups.<sup>11</sup> Consequently, there is still a tremendous need to explore more MOF/CP systems as new TE materials to understand the impacts of the metal ions, their oxidation state, the coordinating functions, the topology of the ligands and the dimensionality of the network on the electric and thermal properties. Based on the record properties of  $\text{Cu}_x\text{BHT}$  and the high stability of copper-thiolate compounds, due to the strong interactions between the thiolate ligands (soft base) and the Cu(I) (soft acid), copper-thiolate CPs are very promising systems as TE materials.<sup>12</sup> Apart from  $\text{Cu}_x\text{BHT}$ , few more conducting copper-monothiolate CPs have been reported, such as  $[\text{Cu}(6\text{-mna})]_n$  ( $5.4 \times 10^{-5} \text{ S.cm}^{-1}$  at RT),<sup>13</sup>  $\{[\text{Cu}_2(6\text{-Hmna})(6\text{-mna})]\cdot\text{NH}_4\}_n$  ( $11 \text{ S.cm}^{-1}$  at RT on single crystal)<sup>14</sup> (6-Hmna = 6-mercaptopyridine-2-thione, 6-mna = 6-mercaptopyridine-2-thione),  $[\text{Cu}(p\text{-SPHOH})]_n$  ( $120 \text{ S.cm}^{-1}$ ),<sup>15</sup> and  $\text{Cu}[\text{Cu}(\text{pdt})_2]$ , (pdt = 2,3-pyrazinedithiolate) ( $6 \cdot 10^{-4} \text{ S.cm}^{-1}$  at RT), pointing out their ability as semiconducting CPs.<sup>16</sup> Here we report, the copper(I)-1,3-benzenedithiolate,  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , a new 2D CP that is made of aligned 1D Cu-S tubes connected together through the bridging ligands. This CP exhibits intrinsic and bulk electrical conductivity of  $1.5 \text{ mS.cm}^{-1}$  under air at RT and a Seebeck coefficient of  $420 \mu\text{V.K}^{-1}$ , a record value among CPs and MOFs, and equivalent to inorganic materials.

$[\text{Cu}_2(1,3\text{-BDT})]_n$  is obtained by mixture of two equivalents of copper(I) chloride with one equivalent of 1,3-benzenedithiol in an acidic aqueous solution, heated in a sealed vial at  $120^\circ\text{C}$ . As shown by the scanning electron microscopy image (Fig. S1), thin needles of few micrometers length are formed. Elemental and thermo-gravimetric analyses lead to the chemical formula

of  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , which confirms the S / Cu(I) = 1 ratio and the charge balance (see ESI and Fig. S2). The infrared spectrum also confirms the coordination of the thiolate functions with the absence of thiol function, usually observed at  $2500 \text{ cm}^{-1}$  (Fig. S3). The powder X-ray diffraction (PXRD) data, collected at the CRISTAL beamline of the SOLEIL synchrotron facility (Gif-Sur-Yvette, France), show good crystallinity of the compound and allow its *ab initio* structural determination. The final Rietveld refinement (Fig. S4) leads to satisfactory model indicator and profile factors (Table S2).  $[\text{Cu}_2(1,3\text{-BDT})]_n$  crystallizes in the monoclinic  $P2_1/c$  space group. Its structure is made of inorganic Cu-S chains, where copper and sulphur atoms are three coordinated, forming distorted  $\text{Cu}_3\text{S}_3$  hexagons packed together along *b* axis (Fig. 1a and b, S5). These Cu-S tubular chains are then connected together through the bridging 1,3-BDT ligands to form a 2D network (Fig. 1c), totally different to that observed for the 3D  $[\text{Ag}_2(1,3\text{-BDT})]_n$  analogue.<sup>17</sup> Such Cu(I)-S tubular chains made of distorted  $\text{Cu}_3\text{S}_3$  hexagons have been already observed in few  $[\text{Cu}(\text{SR})]_n$  CPs with monotopic thiolate ligands, with R = Me,<sup>18</sup> Ph, *p*-PhMe, *p*-PhOMe<sup>19</sup> and *o*-PhCO<sub>2</sub>H.<sup>20</sup> In this 2D CP, the Cu-S distances range from  $2.20(2) \text{ \AA}$  to  $2.26(3) \text{ \AA}$ , the Cu-S-Cu angles from  $75.0(7)^\circ$  to  $125(1)^\circ$  and the S-Cu-S angles from  $109.1(7)^\circ$  to  $127.2(8)^\circ$  (Table S3). However,  $[\text{Cu}_2(1,3\text{-BDT})]_n$  is the first 2D copper-thiolate CP bearing with these 1D Cu-S tubular chains, thus making it possible to combine in one compound, a 1D inorganic chain and a 2D metal- $\pi$ -conjugated path. The Cu-Cu distances of  $2.947(4) \text{ \AA}$  and  $3.797(3) \text{ \AA}$  point out the absence of cuprophilic interactions, that can sometimes be observed in  $d^{10}$  coinage metal compounds.<sup>21</sup> The X-ray photoelectron spectroscopy (XPS) of  $[\text{Cu}_2(1,3\text{-BDT})]_n$  also confirms the oxidation state of Cu(I). Indeed, the 2p-copper core levels exhibit dominant peaks at  $932.8 \text{ eV}$  and  $952.6 \text{ eV}$ , corresponding to Cu(I) 2p<sub>3/2</sub> and Cu(I) 2p<sub>1/2</sub>, respectively (Fig. S6).

The direct band gap obtained from the Tauc plot of the UV-visible absorption spectroscopy curve is  $2.26 \text{ eV}$  (Fig. S7), which is of the same order of magnitude as that of the other reported 1D copper(I)-thiolate CPs ( $2.23\text{--}2.68 \text{ eV}$ ).<sup>19-20</sup> The thermoelectric properties of  $[\text{Cu}_2(1,3\text{-BDT})]_n$  were evaluated from dense pellets with a diameter of  $4 \text{ mm}$  and a thickness of  $1 \text{ mm}$  (Fig S8), prepared at room temperature



**Figure 1.** Structure of  $[\text{Cu}_2(1,3\text{-BDT})]_n$ : a. the Cu-S chain along the *b* axis; b. the two deformed  $\text{Cu}_3\text{S}_3$  hexagons and c. the central projection along the *b* axis. Orange, yellow and gray spheres correspond to copper, sulfur and carbon atoms. Hydrogen atoms were omitted for clarity.

under pressure, with typically, around 40 mg of crushed sample pressed at 0.6 MPa for 15 minutes. PXRD data of these pellets show that the sample remains crystalline and stable under this mechanical pressure, and that no decomposition is observed when it is stored under air for few months (Fig. S8).

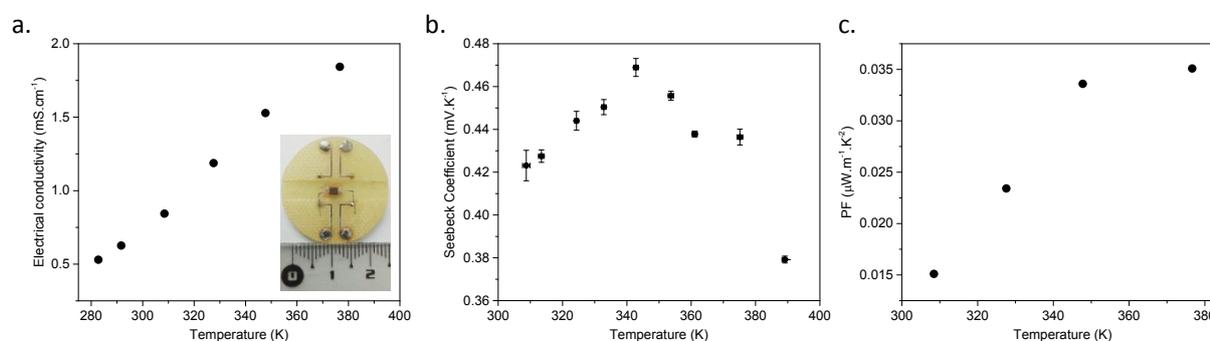
The electrical transport properties were investigated first by four-probe method at the surface of the pellet from RT to 390 K. At RT, the current–voltage measurements on the pellets in the four-probe configuration, show a linear behaviour following the Ohm's law (Fig. S9, which reveals an electrical conductivity of  $8.3 \cdot 10^{-4} \text{ S.cm}^{-1}$ . This modest value is close to the one of the 1D  $[\text{Cu}(\text{SPh})]_n$  analogue, which has an electrical conductivity of  $3.7 \cdot 10^{-4} \text{ S.cm}^{-1}$ , pointing out the implication of the Cu-S chains in the electron transport rather than the aromatic benzene ring.<sup>22</sup>

The variable-temperature measurements with a four-probe configuration using metallic tips applied on the surface of a pressed pellet show that the electrical conductivity increases with the rise in temperature, which is consistent with bulk transport dominated by a thermally activated process (Fig. 2a and S10). Fitting the data with the Arrhenius equation yields an activation energy ( $E_a$ ) of 0.13 eV (Fig. S11).<sup>23</sup> In order to ensure that the electrical conductivity is not a surface effect on the pressed pellet, bulk conductivity also was measured. For that, a corded saw was used to cut the sintered pellet into a rectangular cuboid of  $1.4 \times 2 \times 2.5 \times 4 \text{ mm}$  (Fig. S12) and two electrodes have been welded on the metallised sides of this cuboid to apply a uniform current. The voltage is then measured between differently spaced electrodes, allowing the measurements of the voltage drop as function of distances, in order to extract the resistance and the contact resistance (see insert of Fig. 2a and S13). A low contact resistance of 0.017 k $\Omega$  was determined (two orders of magnitude lower than the one of the sample). The bulk electrical conductivity of the CP measured in this configuration is  $15 \cdot 10^{-4} \text{ S.cm}^{-1}$ , close to the value obtained with the previous surface four-probe experiment, proving that the electrons can move through the bulk pressed pellet prepared without any care of orientation.

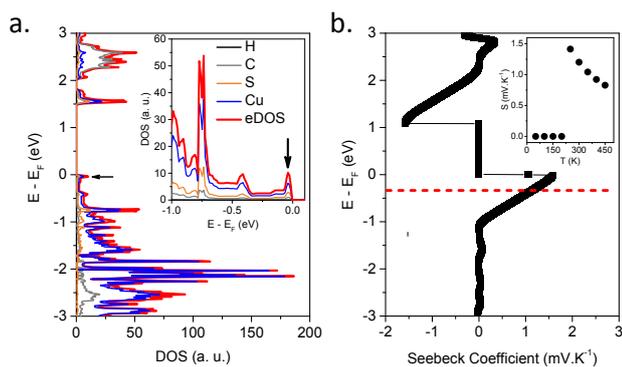
The Seebeck coefficient ( $S$ ) was measured on bulk samples of  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , inserted in a primary vacuum chamber, at different temperatures ranging from RT to 400 K, below the

decomposition temperature of 520 K (Fig. S2). The  $S$  value at RT is positive and high:  $+420 \pm 10 \mu\text{V.K}^{-1}$  (Fig. S14). The positive sign indicates that the dominant charge carriers are holes with a Fermi level close to the valence band. This  $S$  is the highest reported among conducting MOFs, CPs, composite materials and organic polymers (Table S1).<sup>7</sup> Interestingly, the  $S$  increases upon heating up to a certain temperature, around 350 K, from which it decreases reversibly up to 380 K (see Fig. 2b). This unusual existence of temperature maximum will be correlated below to the particularities of the electronic band structure.<sup>24</sup> The power factor (PF), expressed as  $\text{PF} = S^2\sigma$ , is  $0.026 \mu\text{W.m}^{-1}.\text{K}^{-2}$  at RT and reaches  $0.034 \mu\text{W.m}^{-1}.\text{K}^{-2}$  at 349 K. Due to the modest electrical conductivity of  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , this PF is below the best values reported for MOFs and CPs with a record of  $88.2 \mu\text{W.m}^{-1}.\text{K}^{-2}$  for the nanosheets of  $\text{Cu}_x\text{BHT}$  (Table S1).<sup>9b</sup> Nevertheless, as mention earlier, the electrical conductivity of  $[\text{Cu}_2(1,3\text{-BDT})]_n$  is an average over all the directions of the needle-like crystallites constituting the powder, that has been simply pressed to form the pellets. Further studies on aligned crystallites or on a single crystal would allow extracting the intrinsic conductivity of the 1D-CuS chains that might be much higher.

In order to understand the origin of the high Seebeck values and its anomalous temperature dependence in  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , the electronic band structure has been calculated and the Seebeck coefficient using *ab initio* DFT calculations determined. In the optimized  $[\text{Cu}_2(1,3\text{-BDT})]_n$  unit cell, the sulfur atoms are negatively charged, while the copper is in the Cu(I) oxidation state. The electronic band structure along the main symmetry directions is reported in the SI (S15a). A direct band gap of 1.5 eV is extracted from the band structure. This value is underestimated at the current theoretical level, but is sufficient to analyse the nature of the electronic states which mainly contribute to the electrical conductivity and the Seebeck coefficient. Interestingly, the band structure exhibits very flat bands along the directions in momentum space perpendicular to the Cu-S chains, confirming the 1D like character of the electrical conduction. These flattened bands result in peaks when integrated in momentum as shown in the total, eDOS, and projected electron density of states, p-eDOS (see black arrows on Fig. 3a and S15b). Moreover, at the Fermi



**Figure 2.** a. Temperature dependence of the electrical conductivity measured on a  $[\text{Cu}_2(1,3\text{-BDT})]_n$  pressed pellet by surface four-probe configuration and photo of the device to apply the bulk measurement of the electrical conductivity; b. Seebeck coefficient with temperature (errors come from the linear fits of the raw data) and c. the power factor with temperature.



**Figure 3.** a. Calculated density of states (DOS) of  $[\text{Cu}_2(1,3\text{-BDT})]_n$ : projected electronic DOS (p-eDOS) with grey line for C, black for H, orange for S and blue for Cu electrons and the total electronic DOS (eDOS) in red. The insert is a zoom close to the Fermi level and with switched axes. The black arrows show the sharp peaks associated to high Seebeck coefficient. b. Calculated Seebeck coefficient at 300 K. The insert shows its temperature dependence at constant energy represented by the dotted red line on the Figure b.

level, near the top of the valence band, the p-eDOS shows that the dominant character of the propagating electronic states is provided by sulfur and copper orbitals, indicating that there is no conduction between the Cu-S chains, which also confirms that the conduction takes place in the 1D-like Cu-S channels.<sup>24</sup> On the other hand, at the bottom of the conduction band, a more significant contribution of the C-backbone appears, which indicates an inter-chains conduction.

It is well known in the eDOS that sharp features close to the Fermi level result in large values of the Seebeck coefficient.<sup>25</sup> From the transport integrals, the Seebeck is enhanced by sharp and thermally activated electronic singularities in the eDOS. As a direct consequence of the 1D-like character of the valence bands close to the Fermi level, the eDOS of  $[\text{Cu}_2(1,3\text{-BDT})]_n$  exhibits such sharp peaks (see black arrow on Fig. 3a). This explains the large and abrupt changes of the Seebeck coefficient according to energy close to the Fermi level as seen in Fig. 3b, which can also be observed by varying the temperature as shown in the insert of Fig. 3b (and Fig. S16). This qualitatively explains the high positive Seebeck values observed experimentally and also the maximum observed in the temperature dependence of the Seebeck (insert in Fig. 3b). Thus, the reasons for both the high Seebeck values and a maximum value around 350 K are due the strong quasi-1D character of the electronic transport. This can be seen from the *ab-initio* simulated electronic band structures (Fig. S16), the flatness of the bands along the Cu-S chains is responsible for peaks (Van Hove singularities) in the eDOS and the temperature dependence of  $S(T)$ .

## Conclusions

A neutral 2D copper-thiolate CP,  $[\text{Cu}_2(1,3\text{-BDT})]_n$ , has been synthesized under hydrothermal conditions. This CP shows good air stability, due to the strong soft base – soft acid

interactions between the thiolate functions and the copper(I) ions, and it remains crystalline upon pressure for pellet formation. This Cu(I)-thiolate network exhibits at RT a mild conductivity of  $1.5 \text{ mS.cm}^{-1}$  and a record Seebeck coefficient of  $420 \mu\text{V.K}^{-1}$  among CPs, equivalent to inorganic solids. The resulting PF is  $0.03 \mu\text{W.m}^{-1}\text{.K}^{-2}$ . Theoretical studies showed that the conductivity is originating from the 1D inorganic Cu-S chains and explains the high Seebeck coefficient by the presence of a significant peak in the total electronic DOS at the Fermi level, near the top of the valence band. This study highlights that CPs, and more especially copper-thiolates, hold great promises as emerging TE materials. In addition, compared to the highly conducting 2D copper-benzenehexathiolate, which required multistep synthesis of the ligand, 1,3-benzenedithiol is commercially available, reducing considerably the time and cost of the final material production, and allowing formation of large pressed pellets under air. Finally, systematic syntheses of different copper-thiolate CPs and in-depth electronic and theoretical studies have to be done in order to increase the electrical conductivity while keeping a good Seebeck coefficient. Changing ligand functionalities,<sup>12b</sup> mixing sulfur, using selenium or tellurium-based linkers,<sup>26</sup> carrying out metal doping or synthesising heterometallic composition CPs are some paths to improve the electrical conductivity.<sup>27</sup>

## Author Contributions

CA, SH, AA and JLR performed the synthetic experiments and routine characterizations; AM and NG solved the structure; FP was involved in the purification; SW hosted AD; TN performed the DFT, RD and CA carried out the conductivity; OB executed the XPS; SP and AD conceived the ideas, designed the experiments and wrote the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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