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Stanislaw Dzwigaj, Diane Reja, Saremlé Koné-Guira, Antoine Miche, Guylène Costentin, et al.. Cobalt on dealuminated-Si $\beta$  as a catalyst for the oxidative dehydrogenation of propane. Applied Catalysis A : General, 2023, 657, pp.119119. 10.1016/j.apcata.2023.119119 . hal-04193203

**HAL Id: hal-04193203**

**<https://hal.science/hal-04193203>**

Submitted on 1 Sep 2023

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# Cobalt on Dealuminated-Si $\beta$ as a Catalyst for the Oxidative Dehydrogenation of Propane

Stanislaw Dzwigaj, Diane Reja, Saremlé Koné-Guira, Antoine Miche, Guylène

Costentin and Cyril Thomas\*

Sorbonne Université, CNRS, Laboratoire Réactivité de Surface, LRS, F-75005 Paris, France.

\* To whom correspondence should be addressed:

Dr. Cyril Thomas

Sorbonne Université, CNRS, Laboratoire Réactivité de Surface, LRS, UMR CNRS 7197, 4 Place  
Jussieu, Tour 43-53, 3<sup>ème</sup> étage, Case 178, F-75252, Paris, France

e-mail: [cyril.thomas@sorbonne-universite.fr](mailto:cyril.thomas@sorbonne-universite.fr)

Tel: + 33 1 44 27 36 30

Fax: + 33 1 44 27 60 33

## ORCID:

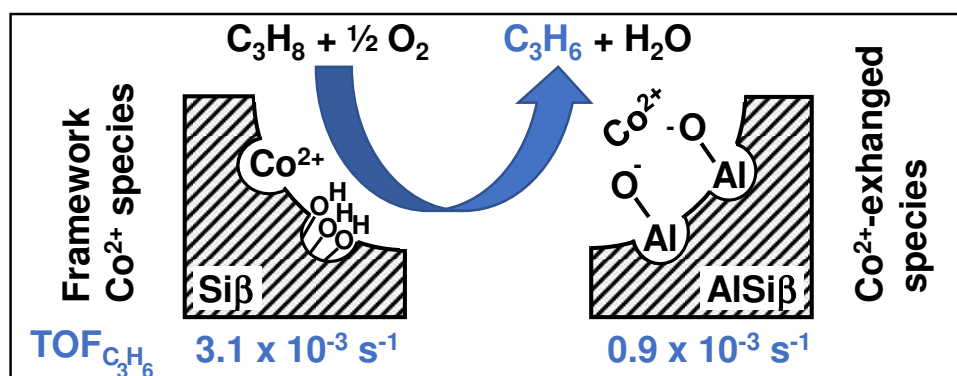
Stanislaw Dzwigaj: [orcid.org/0000-0003-1363-9316](https://orcid.org/0000-0003-1363-9316),

Antoine Miche: [orcid.org/0000-0002-8006-1123](https://orcid.org/0000-0002-8006-1123),

Guylène Costentin: [orcid.org/0000-0003-1559-6890](https://orcid.org/0000-0003-1559-6890),

Cyril Thomas: [orcid.org/0000-0003-4224-6095](https://orcid.org/0000-0003-4224-6095)

## Graphical Abstract:



## Highlights:

- TOF<sub>C<sub>3</sub>H<sub>6</sub></sub> of Co<sup>2+</sup> into Siβ is ~ 4 times higher than that of Co<sup>2+</sup>-exchanged species
- TOF<sub>C<sub>3</sub>H<sub>6</sub></sub> at 400 °C of Co-Siβ is ~ 7 times greater than that of V-Siβ
- Maximum incorporable content of Co as Co-Siβ is dictated by the amount of removable Al

**Abstract:**

The present work aims at investigating the oxidative dehydrogenation of propane (ODHP) catalytic properties of Co species supported on a Si $\beta$  zeolitic support. The Co<sup>2+</sup> species incorporated in the silanol nests created by dealumination of a HAlSi $\beta$  pristine support (Co-Si $\beta$ ) appeared to be about 4 times more active ( $3.1 \times 10^{-3} \text{ s}^{-1}$ ) compared with that in exchange position of HAlSi $\beta$  ( $0.9 \times 10^{-3} \text{ s}^{-1}$ ) for the formation of propene at 400 °C. The turn over frequency of the Co-Si $\beta$  samples are found to be greater than or equal to those reported in most of the earlier studies performed with Co-containing materials to date and to be about 7 times greater than that of V-Si $\beta$  materials ( $0.4 \times 10^{-3} \text{ s}^{-1}$ ) from an earlier study, which illustrates the superiority of Co-Si $\beta$  over V-Si $\beta$  for the ODHP reaction. Maximum incorporable content of Co as Co-Si $\beta$  is dictated by the amount of removable Al.

**Keywords:** Catalysis, ODHP, propene, NO-TPD, XPS, Co-Si $\beta$

## 1. Introduction

Propene ( $C_3H_6$ ) is known as an essential platform chemical molecule for the production of polypropylene (60 %), acrylonitrile (9 %), oxo alcohol (8 %), propylene oxide (7 %) , cumene (6 %) and acrylic acid (4 %) [1]. The industrial production of propene is essentially achieved by steam reforming and fluid catalytic cracking of oil fractions [2-5]. The production of propene via the latter processes does not meet the market demand, however, which is known as the so-called propylene gap [5, 6]. The shale gas resources, comprising substantial amounts of ethane and propane, and the propylene gap renewed the interest in the valorization of propane to propylene. To meet the ever-increasing propene demand, about 110 million metric tons in 2018 [3, 7], the propane direct dehydrogenation (PDH) process has been industrialized [8]. Such a process obviously produces propene with the highest selectivity [7] and needs to be operated at relatively high temperatures, however, due the endothermicity of the dehydrogenation reaction [9], which makes it relatively energy intensive [8]. Another major drawback of the PDH process is related to the deactivation of the catalysts at such elevated temperatures with the deposition of carbonaceous species [8], which remediation requires periodic calcination of these deposits at high temperatures before a reduction step to regenerate the PtSn alloys [10].

Alternative processes are therefore sought for the production of propene in a more sustainable way, such as that of the oxidative dehydrogenation of propane (ODHP) [1-4]. This non-thermodynamically restricted exothermic process is favorably operated at lower temperatures and should be less sensitive to coke deposition because of the presence of oxygen in the reacting feed compared with the PDH process [1-4]. One major drawback of ODHP lies in the lower selectivity in propene compared to PDH due to overoxidation of propane and/or propene by oxygen by side reactions, which severely limits the industrial application of ODHP [2, 8].

Many efforts have been devoted to the study of ODHP catalysts for which metal-oxide based materials have been reported to be most active [4] before boron-containing materials showed promising properties rather recently [11, 12]. Over metal-oxide catalysts, ODHP has been shown to proceed via a Mars and van Krevelen mechanism [13] in which lattice oxygen species of the catalyst is involved in the oxidation process of the propane molecule, with the oxygen vacancies created by water release being replenished by molecular oxygen. The nature of the reacting oxygen species has been suggested to be of prime importance in this reaction for the selectivity in propene. The nucleophilic oxygen species were thus suggested to be much more selective for partial oxidation of propane compared with the electrophilic counterparts that lead to undesired combustion of propane/propene to  $\text{CO}_x$  ( $\text{CO} + \text{CO}_2$ ) [3, 4, 14]. For metal-oxide based catalysts, those including vanadium [15, 16] have been shown to be among the most active in the ODHP reaction [14]. The nature of the supporting oxide [1] and the dispersion of V [3] have been shown to exhibit a drastic influence on the ODHP catalytic performance. Not only the nature of environment of the transition metal is important for ODHP but also the nature of the transition metal in itself. Several studies have reported on the promising properties of Co catalysts for the ODHP reaction indeed [17-30].

Following the pioneering works of Saxton et al. [31], Kim et al. [32] and Kraushaar and Van Hooff [33], Dzwigaj et al. [34] have shown that incorporation of V could be done in the silanol nests [35] created by dealumination of an  $\text{HAlSi}\beta$  pristine support. Later, it was shown that such V-Si $\beta$  catalysts exhibited similar ODHP performance as that reported for vanadium supported on various  $\text{SiO}_2$  supports [36]. Cobalt could also be inserted successfully in the silanol nests by Dzwigaj and co-workers [37-40] to be applied in NO selective catalytic reduction by ethanol [41] and ammonia [42-44],  $\text{N}_2\text{O}$  decomposition [42], dehydrochlorination [45] and Fischer-Tropsch

synthesis [46, 47]. As far as we know, the performance of Co-Si $\beta$  samples has not been reported in the ODHP reaction. The present work therefore aims at investigating the catalytic properties of Co species located in various environments of a  $\beta$  zeolitic support and to compare the performance of such species with earlier literature data reported on Co-based catalysts. In particular, the influence of the pH of the impregnating aqueous solutions was studied on the incorporation of Co<sup>2+</sup> species in Si $\beta$ .

## **2. Experimental**

### **2.1. Catalyst preparation**

The synthesis of the Co<sub>x</sub>-Si $\beta$  series of samples, which is schematically described in ref. [40], was done via a slightly modified two-step post-synthesis method developed earlier for the preparation of V-Si $\beta$  materials [34, 48].

#### **2.1.1. Preparation of the HAlSi $\beta$ and Si $\beta$ supports**

Briefly, the tetraethyl-ammonium  $\beta$  pristine zeolite (TEA $\beta$ , SINOPEC-RIPP, China, Si/Al ~ 14) was firstly calcined in a muffle furnace at 550 °C for 15 h to remove the TEA template and to produce the protonic form of the zeolite material (HAlSi $\beta$ ). HAlSi $\beta$  (12 g) was then dealuminated in a nitric-acid solution (1 040 mL of concentrated HNO<sub>3</sub> (Aldrich, 68 wt% HNO<sub>3</sub>) with 160 mL of distilled water resulting in a nitric acid concentration of 13 mol/L) at 80 °C for 4 h. The suspension was then cooled to RT and the solid was filtered before being washed with distilled water. Finally, the dealuminated material (Si $\beta$ , Si/Al ~ 400) was dried at 80 °C overnight in a laboratory oven.

### **2.1.2. Deposition of Co on Si $\beta$**

Deposition of Co was carried out by contacting 2 g of Si $\beta$  with 200 mL of Co(NO<sub>3</sub>)<sub>2</sub> (Aldrich) aqueous solutions (from about 1.6 to 17.4 x 10<sup>-3</sup> mol Co/L), either at pH 2.5 or 9.0, for 24 h under continuous stirring. The pH of these solutions was adjusted by adding drops of either concentrated nitric acid or ammonia solutions (Aldrich). Such a procedure led to Co loadings ranging from about 1 to 9 wt%. It must be stressed that the addition of the Si $\beta$  support in the Co(NO<sub>3</sub>)<sub>2</sub> aqueous solutions at pH 9.0 resulted in a rapid decrease in the pH of the suspensions to about pH 3.0. In that case, the pH of the suspensions was then adjusted to pH 9.0 by addition of concentrated ammonia. The suspensions were then transferred in a rotary evaporator and water was evaporated under vacuum at 70 °C for about 2 h. As the ODHP reaction was performed up to 500 °C (Section 2.3.), the samples were calcined in a muffle furnace at 500 °C (100 °C/h) for 3 h. These series of samples will be referred to as Co<sub>x</sub>-Si $\beta$ -pH-2.5 and Co<sub>x</sub>-Si $\beta$ -pH-9.0, respectively, where *x* represents the Co loading (wt%).

### **2.1.3. Deposition of Co on HAlSi $\beta$**

For comparison, a sample including about 1.8 wt% Co was also prepared by wet impregnation of the HAlSi $\beta$  support (Co<sub>1.8</sub>-HAlSi $\beta$ ) by an aqueous solution of Co(NO<sub>3</sub>)<sub>2</sub> following the same procedure as that described for the two previous series of samples but at pH 3.0.

## **2.2. Catalyst characterization**

ICP-OES, XPS and H<sub>2</sub>-TPR investigations were performed on the samples submitted to NO-TPD non-destructive experiments.



The chemical composition of the samples was ascertained by ICP-OES (Inductively Coupled Plasma Optical Emission Spectroscopy) using a 5100 SVDV spectrometer (Agilent) at the Plateforme ALIPP6 Laboratoire d'Analyse Géo chimie Élémentaire (Sorbonne Université). The samples (50 mg) were firstly digested in a mixture of 1 mL of concentrated HF (Normatom, 47 wt%) and 1 mL of concentrated HNO<sub>3</sub> (68 wt%) at 80 °C for 2 h. Excess HF was then “neutralized” by 50 mL of an aqueous solution of H<sub>3</sub>BO<sub>3</sub> (20 g/L) by complexation of the fluorides. The obtained solution was finally diluted with a diluted HNO<sub>3</sub> aqueous solution if necessary. Appropriate calibration mixtures were also used.

XP spectra were collected on an Omicron Scienta (Argus) X-ray photoelectron spectrometer using a monochromated Al K $\alpha$  ( $h\nu = 1486.6$  eV) radiation source having a 280 W electron beam power. Charge compensation was carried out using an electron beam having an energy of 1 eV and an emission current of 5  $\mu$ A. The emission of photoelectrons from the sample was analyzed at a takeoff angle of 45° under ultrahigh vacuum conditions ( $1 \times 10^{-7}$  Pa). XP spectra were collected at a pass energy of 20 eV for C 1s, O 1s, Si 2p, Al 2p and Co 2p core XPS levels. After data collection, because of the known difficulties in referencing the binding energies (BEs) to the C 1s line [49], the BEs were calibrated with respect to the BE of the Si 2p peak (FWHM  $\sim$  1.8 eV) at 103.3 eV [44, 50, 51]. All of the BEs reported in this work were measured within an accuracy of  $\pm 0.2$  eV. The peak areas were determined after subtraction of a Tougaard background. The atomic ratio calculations were performed after normalization using Scofield factors [52]. Spectrum processing was carried out using the Casa XPS software package and Origin 7.1 (Origin Lab Corporation). The Co<sub>3.0</sub>Si $\beta$ -pH-2.5 and Co<sub>3.1</sub>Si $\beta$ -pH-9.0 samples were also analyzed by XPS after ODHP reaction.

Temperature-Programmed Reduction by hydrogen (H<sub>2</sub>-TPR) experiments were performed by using a Micromeretics 2920 equipment. 40 to 50 mg of the sample was introduced in the U-type reactor on a plug of quartz wool and submitted to oxidation under 25 mL/min O<sub>2</sub>(5%)/He from RT to 500 °C (10 °C/min) and maintained at 500 °C for 1 h. It was found that the samples did not exhibit any oxidation contribution (**Figure S1**), indicating that previous NO-TPD experiments did not lead to the reduction of the Co species. The samples were then cooled to RT under the O<sub>2</sub>(5%)/He flow. After purging by Ar (25 mL/min), the samples were submitted to H<sub>2</sub>-TPR under 25 mL/min H<sub>2</sub>(5%)/He from RT to 900 °C (10 °C/min) and left at this temperature for 2 h. Water release in H<sub>2</sub>-TPR experiments was trapped with an isopropanol-liquid N<sub>2</sub> mixture prior to the TCD detector.

NO Temperature-Programmed Desorption (NO-TPD) experiments were carried out in a U-shape quartz reactor (10 mm i.d.). The samples (about 0.12 g) were held on a plug of quartz wool and the temperature was controlled by a Eurotherm 2408 temperature controller using a K-type thermocouple. Reactant gases (Air Liquide) were used as received and fed from independent gas cylinders by means of mass flow controllers (Brooks 5850TR) with a total flow rate of 230 mL<sub>NTP</sub>/min. The reactor outlet was continuously monitored by chemiluminescence NO<sub>x</sub> analyzers (Thermo Environmental Instruments 42C-HL or Eco Physics CLD 700 AL) that allowed the simultaneous detection of NO and NO<sub>2</sub> (NO<sub>x</sub> = NO + NO<sub>2</sub>). Prior to the NO-TPD experiments, the samples were annealed in situ in He or Ar at 500 °C (3 °C/min) for 2 h under a flow rate of 100 mL<sub>NTP</sub>/min. Following the in-situ annealing treatment, the samples were exposed to the adsorption mixture that consisted of 400 ppm NO diluted in He or Ar. The samples were exposed to the adsorption mixture at RT for about 2 h until the outlet NO<sub>x</sub> readout was equivalent to the inlet NO<sub>x</sub> readout. The latter parameter was set to ensure that each sample was saturated under the present experimental conditions. Before the NO-TPD step, the samples were flushed in He or Ar at RT for

about 1 h to remove weakly chemisorbed species until the NO and NO<sub>2</sub> concentrations detected at the outlet of the reactor were found to be negligible. NO-TPD experiments were carried out from RT to about 500 °C at a heating rate of 3 °C/min under a flow of 230 mL<sub>NTP</sub>/min of He or Ar. It was checked on two aliquots of Co<sub>1.0</sub>Siβ-pH-2.5 that both NO<sub>x</sub> analyzers provided similar NO<sub>x</sub>-TPD profiles (**Figure 1**) and NO<sub>x</sub> uptakes (105 and 102 μmol/g). It must be noted that only traces of NO<sub>2</sub> were detected over the Co<sub>x</sub>-Siβ-pH-2.5 and Co<sub>x</sub>-Siβ-pH-9.0 series, whereas significant quantities of NO<sub>2</sub> were recorded over HAlSiβ and Co<sub>1.8</sub>-HAlSiβ, the origin of which may be attributed to NO disproportionation ( $3 \text{ NO} = \text{N}_2\text{O} + \text{NO}_2$ ) [53].

### 2.3. ODHP catalytic experiments

Prior to the catalytic runs, the Co<sub>x</sub>-Siβ-pH-2.5, Co<sub>x</sub>-Siβ-pH-9.0 and Co<sub>1.8</sub>-HAlSiβ samples (~ 20 μmol of Co) diluted in 0.6 g of α-Al<sub>2</sub>O<sub>3</sub> (125-200 μm, Aldrich) were calcined in O<sub>2</sub> (2.2%)–He (65.1 mL<sub>NTP</sub>/min) at 500 °C (2 °C/min) for 3 h unless specified otherwise. The temperature was then decreased to 400 °C. The ODHP experiments were carried out in a U-type quartz reactor (10 mm i.d.). The samples were held on a plug of quartz wool, and the temperature was controlled by a Eurotherm 2408 temperature controller using a K type thermocouple. Reactant gases (Air Liquide, used as received) were fed from independent mass flow controllers (Brooks 5850TR and Bronkhorst F-201CV-200-RAD-11-V). Typically, the composition of the C<sub>3</sub>H<sub>8</sub>-O<sub>2</sub>-He (4.3–1.0–47.3) reaction mixture was: 8.2 % C<sub>3</sub>H<sub>8</sub> and 1.9 % O<sub>2</sub> in He with a total gas flow of 74.1 mL<sub>NTP</sub>/min. The catalyst was contacted with the reaction mixture at 400 °C and the temperature was increased by steps of 25 °C (5 °C/min) up to 500 °C for about 1 h 30 at each reaction step at atmospheric pressure. After reaction at 500 °C for 1 h 30, the catalyst was cooled to 400 °C under the reaction mixture. Analysis of the composition of the reactor outflow was performed on-line with a μ-GC

(Varian, CP4900) using three different columns running in parallel. The 5A molecular sieve column (80 °C, 150 kPa He, 200 ms injection time, 30 s backflush time) was used to separate H<sub>2</sub>, O<sub>2</sub>, CH<sub>4</sub> and CO. The poraplot Q column (100 °C, 150 kPa He, 50 ms injection time, 6 s backflush time), was used to separate CO<sub>2</sub>, C<sub>2</sub>H<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, H<sub>2</sub>O, C<sub>3</sub>H<sub>6</sub> and C<sub>3</sub>H<sub>8</sub>. The CP-SIL5 column (45 °C, 150 kPa He, 50 ms injection time) was used to separate C<sub>3</sub>H<sub>6</sub>O compounds (acetone + propenal) from C<sub>3</sub>H<sub>4</sub>O (acrolein).

The conversion of C<sub>3</sub>H<sub>8</sub> (X C<sub>3</sub>H<sub>8</sub>) and the product selectivities (S A<sub>i</sub>) were calculated as follows:

$$X \text{ C}_3\text{H}_8 (\%) = (\sum n_i [A_i] / (3 \times [\text{C}_3\text{H}_8]_{\text{inlet}})) \times 100$$

$$S \text{ A}_i (\%) = (n_i [A_i] / \sum n_i [A_i]) \times 100$$

where n<sub>i</sub>, [A<sub>i</sub>] and [C<sub>3</sub>H<sub>8</sub>]<sub>inlet</sub> are the number of carbon atoms in the A<sub>i</sub> products, the concentrations of the A<sub>i</sub> products, and the inlet concentration of propane, respectively.

For comparison with earlier literature data, the formation of C<sub>3</sub>H<sub>6</sub> in the ODHP catalytic reaction was calculated as turnover frequencies (TOFs) [54] normalized by the total number of Co atoms such as:

$$\text{TOF (s}^{-1}\text{)} = (([\text{C}_3\text{H}_6] \times F) / (100 \times V_m)) / (sw \times \% \text{Co}) / (100 \times M_{\text{Cobalt}}),$$

Where [C<sub>3</sub>H<sub>6</sub>], F, V<sub>m</sub>, sw, %Co and M<sub>Cobalt</sub> are the concentration in C<sub>3</sub>H<sub>6</sub> at the outlet of the reactor (%), the total flow rate (mL s<sup>-1</sup>) at 0 °C and 101 325 Pa, the molar volume (22414 mL mol<sup>-1</sup>) at 0 °C and 101 325 Pa, the sample weight (g), the Co content (wt%) and the atomic weight of Co (58.933 g mol<sup>-1</sup>), respectively.

Under the present experimental conditions, it was checked that the reactor loaded with 0.6 g of α-Al<sub>2</sub>O<sub>3</sub> did not contribute to the transformation of C<sub>3</sub>H<sub>8</sub> to a significant extent at 400 °C (**Figure S2**)

and that both  $C_3H_8$  and  $O_2$  conversions measured at 400 °C varied proportionally to the amount of sample loaded in the reactor, indicating the absence of external limitations. The carbon and oxygen balances were found to be above 98 and 90 %, respectively.

### 3. Results and discussion

#### 3.1. Speciation of Co in the impregnating solutions

The speciation of Co in the aqueous medium was computed [55] as a function of the pH of the aqueous medium (by addition of  $HNO_3$  or  $NH_3$ ) for the lowest (**Figure 2a**) and the greatest (**Figure 2b**) Co concentrations used for the preparation of the materials. As illustrated in **Figure 2a**, Co remains as  $Co(H_2O)_6^{2+}$  species up to pH 8.0. At higher pH, the precipitation of  $Co(OH)_2$  occurs in agreement with an earlier literature report [56].  $Co(OH)_2$  appears to be the dominant species up to about pH 11 before the appearance of various Co amine complexes at higher pH. A comparable trend is observed for the highest Co concentration (**Figure 2b**) with a shift of the  $Co(OH)_2$  precipitation frontier to lower pH, however. Overall, it can be deduced from **Figure 2** that Co should be essentially present as  $Co(H_2O)_6^{2+}$  at pH 2.5, whereas it should precipitate to  $Co(OH)_2$  at pH 9.0 for all of the Co concentrations used in the present work for the preparation of the materials.

#### 3.2. Characterization of the samples

XRD,  $^1H$ ,  $^{27}Al$  and  $^{29}Si$  MAS NMR, FTIR and DR-UV-vis complementary data recorded on selected samples can be found in an another article [40]. Briefly, it was found that dealumination of the HAlSi $\beta$  pristine support led to a contraction of the zeolite structure (XRD), a significant

decrease in the intensity of the  $^{27}\text{Al}$  NMR signals and an increase in that of the  $^1\text{H}$  NMR signals, and an increase in the intensity of the OH IR contributions in agreement with earlier studies [34, 36-47]. These data corroborated the extraction of the framework Al species under strongly acidic conditions and the corresponding formation of silanol nests in the dealuminated Si $\beta$  support. The addition of Co to the dealuminated Si $\beta$  support led to an expansion of the zeolite structure (XRD) [37-47] and a pronounced decrease in the intensity of the  $^1\text{H}$  NMR signals [47] and the OH IR contributions [37-40, 42, 47], attesting for the incorporation of  $\text{Co}^{2+}$  ions into the Si $\beta$  framework and the formation of various Co species as indicated by DR-UV-vis spectroscopy [37, 40-44, 57].

### 3.2.1. Chemical composition ascertained by ICP-OES

The chemical composition of the samples deduced from ICP-OES analyses (**Table 1**) shows that the amount of Co deposited agrees reasonably with that targeted. The Si/Al ratio of the pristine support HAlSi $\beta$  was found to be about 14 in agreement with the expected value (Section 2.1.), whereas those of the dealuminated Si $\beta$  and  $\text{Co}_x\text{-Si}\beta$  samples increased to 380 and  $500 \pm 40$ , respectively. The lower Si/Al ratio of the Si $\beta$  support (380) compared to those of the  $\text{Co}_x\text{-Si}\beta$  materials ( $500 \pm 40$ ) may be attributed to further dealumination of the Si $\beta$  support in the course of the deposition step of Co that was performed at low and high pH for 24 h (Section 2.1.). Dealumination of the pristine HAlSi $\beta$  support has been reported to result in the creation of silanol nests in the Si $\beta$  structure [34, 36-39].

### 3.2.2. XPS

The binding energies (BEs) of the O 1s, C 1s, Al 2p and Co 2p $_{3/2}$  core levels of the samples are summarized in **Table 2**. The narrow peak (FWHM  $\sim 1.7$  eV) of the O 1s core level with a fairly

constant BE ( $\sim 532.5$  eV) can be attributed to  $O^{2-}$  species from the Si $\beta$  crystalline network [51, 58, 59].

The presence of adventitious carbon was also detected on all samples at 283.8-284.9 eV ( $13 \pm 1$  at. %). The fact that the BE of C 1s core level varied by about 1 eV may be attributed to imperfect compensation of the local charging of the non-conductive materials despite the use of a flood gun, which provides support for using the Si 2p (**Figure S3**) as a reference rather than adventitious carbon [49].

Al 2p peaks (74.5 eV, FWHM  $\sim 2.2$  eV) attributed to  $Al^{3+}$  species [51, 59] could only be observed on the non-dealuminated HAlSi $\beta$  and  $Co_{1.8}$ -HAlSi $\beta$  samples in agreement with other XPS measurements [40]. The surface Si/Al at. ratio of the non-dealuminated samples (**Table 2**) was found to be in good agreement with the bulk Si/Al ratio determined by ICP-OES (**Table 1**), indicating that Al was homogeneously distributed in the HAlSi $\beta$  pristine material. Given that residual amounts of Al could be measured by ICP-OES on the dealuminated Si $\beta$  samples (**Table 1**), the absence of Al 2p signal on these samples may be assigned to the presence of residual Al atoms in the core of the dealuminated Si $\beta$  grains whose Al 2p core level signal would be screened by the silicon atoms of the outer layers. This suggests that dealumination of the pristine HAlSi $\beta$  support was easier to achieve on the outer layers of the zeolite grains rather than in their cores. The presence of residual Al found by ICP-OES in the dealuminated Si $\beta$  samples appears to be consistent with the presence of an  $^{27}Al$  MAS NMR peak at 51-54 ppm [40, 47] characteristic of tetrahedrally coordinated  $Al^{3+}$  species in the Si $\beta$  zeolite framework.

The Co 2p core level showed essentially the characteristic spin-orbit doublet (Co 2p $_{1/2}$  at ca. 797 eV and Co 2p $_{3/2}$  at ca. 781 eV, **Table 2**) with shake-up satellites at about 803 and 786 eV (**Figure S3**), indicating the presence of  $Co^{2+}$  species in the prepared samples. Whereas the BE of

the Co 2p<sub>3/2</sub> signal remained essentially constant and close to 781.3 eV on the Co<sub>x</sub>-Siβ series prepared at pH 2.5 and 9.0 (**Table 2**), it is of interest to notice that the BE of the Co 2p<sub>3/2</sub> signal of Co<sub>1.8</sub>-HAlSiβ appeared to be recorded at a substantially higher BE (782.5 eV, **Table 2**). In agreement with the earlier study of Boix and Fierro, performed over various zeolitic matrixes [60], such an elevated BE is consistent with the presence of Co<sup>2+</sup> species acting as charge compensation cations on the non-dealuminated pristine support. The Co 2p<sub>3/2</sub>/Si 2p surface at. ratios are plotted as a function of the bulk Co/Si at. ratios deduced from ICP-OES in **Figure 3**. This figure firstly shows that the Co 2p<sub>3/2</sub>/Si 2p surface at. ratios are much smaller than the bulk ones over all of the investigated samples in line with an earlier study [44]. Such a discrepancy may be accounted for by a screening effect of the Co signal by the Si atoms from the β zeolite network. In addition, it can be seen that the Co 2p<sub>3/2</sub>/Si 2p surface at. ratio varies reasonably linearly with respect to the bulk one over the whole range of compositions of the investigated samples, which indicates that the dispersion state of the Co atoms remained essentially comparable over the various series of samples from an XPS viewpoint.

### 3.2.3. H<sub>2</sub>-TPR

The H<sub>2</sub>-TPR profiles on both series of samples are shown in **Figure 4**. The Co<sub>x</sub>-Siβ-pH-2.5 samples showed essentially a reduction peak at high temperature at about 800 °C (**Figure 4a**). A shoulder also appeared at about 730 °C for Co<sub>5.6</sub>-Siβ-pH-2.5 (arrow in **Figure 4a**) and the reduction of the samples including a Co loading greater than or equal to 1.9 wt% was found to begin at lower temperature (~ 700 °C) compared to that of Co<sub>1.1</sub>-Siβ-pH-2.5 (~ 770 °C, **Figure 4a**). The Co<sub>x</sub>-Siβ-pH-9.0 samples showed a reduction peak at about 800 °C as well (**Figure 4b**). For the most highly loaded samples (Co content greater than or equal to 4.3 wt%), a reduction peak is also observed at



about 330 °C. As also reported in [40], the Co<sub>1.8</sub>-HAlSiβ sample did not exhibit any H<sub>2</sub> reduction feature (**Figure S4**), indicating that the Co<sup>2+</sup> species on charge compensation sites did not reduce below 900 °C [45, 61-63]. On the Co<sub>x</sub>-Siβ series, the high-temperature peak can therefore be assigned to the reduction of framework pseudo-tetrahedral Co(II) species in strong interaction with the Siβ support [37-42], whereas that at about 330 °C can be assigned to the reduction of Co<sub>3</sub>O<sub>4</sub> crystallites. Yet these Co<sub>3</sub>O<sub>4</sub> crystallites could not be detected by XRD on Co<sub>4.3</sub>Siβ-pH-9 [40] likely because of being too small and/or in low amount [61]. The consumption of H<sub>2</sub> in the H<sub>2</sub>-TPR experiments amounted to be about 72 ± 21 % of the Co content (**Figure 4**), as already reported earlier on similar samples [42]. At the lower contents of Co, the high-temperature reduction peak appears to be incomplete, in other words the signal does not return to baseline, and this may be at the origin of an underestimation of the amount of Co reduced in the H<sub>2</sub>-TPR experiments.

#### 3.2.4. NO<sub>x</sub>-TPD

Earlier studies have reported on the interaction of NO with Co<sup>2+</sup> species on various zeolitic materials and the use of temperature-programmed desorption experiments as a tool to characterize these materials [43, 64-66]. Under the present experimental conditions, it was found that the dealuminated support (Siβ) did not chemisorb NO (**Figure S5**) in agreement with previous studies performed on various silica materials [67-69]. In contrast, the HAlSiβ pristine carrier exhibited two desorption peaks at 100 and 320 °C (**Figure 5a**) in agreement with the earlier work of Dzwigaj and co-workers [43]. The greater NO uptake of HAlSiβ (65 μmol/g), determined from the profile shown in **Figure 5a** (purple trace), compared to that reported previously (8 μmol/g, [43]) can be attributed to differences in adsorption temperatures [66]. The marked difference in NO uptake between Siβ and HAlSiβ demonstrates that the presence of Al is instrumental in the chemisorption of NO on a

Co-free sample and makes more complicated the estimation of the NO uptake by the  $\text{Co}^{2+}$  species, and the corresponding NO/Co ratio, for  $\text{Co}_{0.8}\text{HAlSi}\beta$ . **Figure 5a** shows that the addition of Co to HAlSi $\beta$  leads to a significant increase in the amount of NO desorbed (195  $\mu\text{mol/g}$ , gray trace) compared to that of the pristine HAlSi $\beta$  support (65  $\mu\text{mol/g}$ , purple trace). This indicates that NO also interacted with the  $\text{Co}^{2+}$  exchanged species of  $\text{Co}_{0.8}\text{-HAlSi}\beta$  to a significant extent. The chemical composition of  $\text{Co}_{0.8}\text{-HAlSi}\beta$  (**Table 1**) indicates that about 63 % of the protons have been exchanged by the  $\text{Co}^{2+}$  species ( $\text{Co/Al} \sim 0.31$ ) on such sample. Considering this together with (i) the amount of NO desorbed from this sample (195  $\mu\text{mol/g}$ ) and that from the pristine HAlSi $\beta$  support (65  $\mu\text{mol/g}$ ) and (ii) assuming that  $\text{Co}^{2+}$ -charge-compensated  $\text{AlO}^-$  sites did no longer contribute to NO desorption from the HAlSi $\beta$  support, the NO/Co molar ratio can be estimated to be about 0.56 for  $\text{Co}_{0.8}\text{-HAlSi}\beta$ .

Finally, the NO-TPD profile recorded on  $\text{Co}_{0.8}\text{-HAlSi}\beta$  (**Figure 5a**, grey trace) does not resemble any of the TPD profiles described in earlier studies [64-66, 70]. This may be attributed to differences in the TPD experimental conditions and/or to the nature of the studied samples. In addition, small amounts of  $\text{NO}_2$  were also observed in the course of the NO-TPD experiments performed on HAlSi $\beta$  and  $\text{Co}_{0.8}\text{-HAlSi}\beta$  at 130 and 220  $^\circ\text{C}$ , respectively. The formation of  $\text{NO}_2$  may be due to NO disproportionation ( $3 \text{NO} = \text{NO}_2 + \text{N}_2\text{O}$ ), as suggested previously by Addison and Barrer on Na-zeolites [53].

Regarding the Co samples prepared with the dealuminated support (**Figure 5b,c**,  $\text{Co}_x\text{-Si}\beta$ -pH-2.5 and  $\text{Co}_x\text{-Si}\beta$ -pH-9.0 series), the desorbed species can be attributed exclusively to NO interacting with the  $\text{Co}^{2+}$  species, as it was found that the Si $\beta$  bare support did not chemisorb NO (**Figure S5**). The TPD profiles recorded on both series showed a main NO desorption peak at about

80 °C and two additional contributions of much weaker intensity at about 200 and 330 °C. On a comparable sample (Co(PS)SiBEA), Dzwigaj and co-workers [43] reported only one NO desorption peak at 190 °C, which is not inconsistent with the NO-TPD profiles reported in the present work. Firstly, it must be highlighted that the adsorption step on Co(PS)SiBEA was carried out at 70 °C [43], thus at a temperature close to that for which the main desorption peak was observed in the present study (**Figure 5b,c**), which would explain the absence of such contribution in the work of Dzwigaj and co-workers [43]. Secondly, the contribution of weak intensity observed at 200 °C (**Figure 5b,c**) is consistent with that observed earlier at 190 °C by Dzwigaj and co-workers [43]. Thirdly, given that the contribution at 330 °C is generally of weaker intensity than that at 200 °C (**Figure 5b,c**), such a contribution may have remained below detection limits in the work of Dzwigaj and co-workers [43]. Definitive assignment of the observed NO contributions (**Figure 5b,c**) remains uncertain. Earlier studies performed on similar Co-Si $\beta$  materials have shown essentially the formation of dinitrosyl complexes with NO adsorption followed by FTIR [38, 43]. In particular, Mihaylova et al. have reported that two different Co<sup>2+</sup>(NO)<sub>2</sub> dinitrosyls were formed on Co-Si $\beta$  [38]. Over an exchanged Co-ZSM-5 material, Zhang et al. have showed that the decomposition temperatures of two Co dinitrosyl complexes was strongly dependent on the nature of the dinitrosyl complex [70]. The multiple NO peaks observed in the NO-TPD experiments (**Figure 5b,c**) may therefore be attributed to the decomposition of various dinitrosyls, and the potential decomposition of mononitrosyls [38] in the Co<sub>x</sub>-Si $\beta$  materials. Additional work would obviously be needed to ascertain such assumptions.

The NO uptakes deduced from the NO-TPD profiles (**Figure 5b,c**) are displayed in **Figure 5d**. Overall, it can be seen that the NO uptakes are similar on both series of Co<sub>x</sub>-Si $\beta$  of samples and therefore that the pH at which Co was deposited did not exhibit a significant influence on the

dispersion of the Co species for comparable Co loadings. The NO/Co ratios estimated from the amounts of NO desorbed are plotted in **Figure 5e**. The NO/Co ratios of the most lowly loaded samples of both series (Co<sub>0.1</sub>-Siβ-pH-2.5: NO/Co = 0.56 and Co<sub>0.9</sub>-Siβ-pH-9.0: NO/Co = 0.43), for which the highest dispersion of the Co<sup>2+</sup> species may be assumed, are found to be comparable to that found for Co<sup>2+</sup> cations on exchanged position of Co<sub>1.8</sub>-HAlSiβ (NO/Co = 0.56). This suggests that the NO adsorption properties of the Co<sup>2+</sup> species supported on the dealuminated Siβ support should not be affected to a significant extent compared to those of the Co<sup>2+</sup> cations in exchanged position present in Co<sub>1.8</sub>-HAlSiβ. The fact that the NO/Co ratios are found to be much lower than that expected with the formation of dinitrosyls [38], which still remains unclear at the present stage, therefore prevents the use of the NO-TPD data for a quantitative determination of the Co dispersion on the samples. Yet these data can be used to compare the dispersion of Co in the studied materials from a qualitative viewpoint. Within each series of samples, namely Co<sub>x</sub>-Siβ-pH-2.5 or Co<sub>x</sub>-Siβ-pH-9.0, **Figure 5e** shows that the NO/Co ratio decreases as the Co content increases, as also reported by Gutierrez et al. [65] on various Co/mordenite materials. The decrease in NO/Co ratio (**Figure 5e**) can be attributed to a decrease in the dispersion of the Co species as the Co content increases with the formation of Co<sub>x</sub>O<sub>y</sub> species [65]. Overall, the NO uptakes obtained in the present study are found to be lower than those reported on Co-zeolite materials in previous studies [64-66, 70]. This probably relies both on differences in the studied materials and the experimental conditions under which the NO-TPD experiments were performed. In these earlier studies, the NO-TPD experiments have been carried out on various zeolite supports such as ZSM5 [64, 70], Y [64, 66, 70], Mordenite [64, 65, 70] and β [66]. Li and Armor have reported that the nature of the zeolite support was decisive for the chemisorption of NO when exchanged by Co<sup>2+</sup> ions [64]. In addition, part of the NO uptakes has been recorded on H-zeolite supports [64-66], whereas it was shown in

the present study that the presence of protons in the zeolite support increased the NO uptake to a significant extent (**Figure 5a**). Finally, the NO partial pressures used in these earlier studies were much higher (1000 – 2000 ppm NO) than that used in the present study (400 ppm), and this parameter has been suggested to influence the NO/Co stoichiometry [64].

Overall, the characterization of the  $\text{Co}_x\text{-Si}\beta$  series of samples performed by  $\text{H}_2$ -TPR (Section 3.2.3.) and NO-TPD (Section 3.2.4.) seems to indicate that Co(II) species (XPS, Section 3.2.2.) may be involved in various environments depending on the Co loading. Yet the quantification of the various Co species has been reported to be extremely challenging in the case of  $\text{Co}^{2+}$ -exchanged HAlSi $\beta$  materials [71].

### 3.3. ODHP catalytic results

Prior to evaluation of the Co samples, the thermal reaction between  $\text{C}_3\text{H}_8$  and  $\text{O}_2$  was probed in a reactor loaded with 600 mg of  $\alpha\text{-Al}_2\text{O}_3$  (**Figure S2**).  $\text{C}_3\text{H}_8$  conversion, increasing from 0.1 to 1.0 % from 400 to 500 °C, was found to be mainly due to the complete oxidation of propane to carbon oxides ( $\text{CO}_x = \text{CO} + \text{CO}_2$ ) with a selectivity of about  $80 \pm 3$  % in the latter by-products. Traces of methane, ethane, ethylene, acetone/acrolein and acrolein were observed above 450 °C, whereas detectable amounts of  $\text{H}_2$  were recorded above 475 °C. The formation of  $\text{C}_3\text{H}_6$  was also recorded and found to increase from 12 to about 220 ppm from 400 to 500 °C, resulting in a  $\text{C}_3\text{H}_6$  selectivity of about  $14 \pm 2$  %. As the thermally assisted formation of  $\text{C}_3\text{H}_6$  was found to be limited at 400 °C, comparison of the catalytic performance of the various samples was made at such a temperature. Finally, the conversion of  $\text{O}_2$  increased from 2 to 22 % as the reaction temperature increased from 400 to 500 °C.

Bian et al. have reported recently that highly dispersed tetrahedrally coordinated Co(II) species, stabilized on a mesoporous silica, efficiently catalyzed the direct dehydrogenation of C<sub>3</sub>H<sub>8</sub> (PDH) at 600 °C [72]. Although PDH has been reported to remain thermodynamically disfavored below about 550 °C [56], Davis and co-workers [73], and Wang et al. [74] showed that such a reaction could be catalyzed at 500 and 550 °C by Co(II) single-sites supported on N-doped carbon and ordered SiO<sub>2</sub> channels, respectively. By removing O<sub>2</sub> from the ODHP reacting feed at 400 °C, it was verified that PDH did not contribute to the production of C<sub>3</sub>H<sub>6</sub> to a significant extent under our conditions on Co<sub>3.0</sub>-Siβ-pH-2.5 and Co<sub>3.1</sub>-Siβ-pH-9.0 (**Figure S6**), however.

O<sub>2</sub>, CO<sub>x</sub> and C<sub>3</sub>H<sub>6</sub> concentrations, and C<sub>3</sub>H<sub>8</sub> conversion are displayed in **Figure 6** as a function of time on stream and reaction temperature for the catalysts loaded with about 2 wt % Co. As expected, C<sub>3</sub>H<sub>8</sub> conversion, CO<sub>x</sub> and C<sub>3</sub>H<sub>6</sub> concentrations increased, whereas O<sub>2</sub> concentration decreased as the reaction temperature increased for all of the samples. **Figure 6** shows that the catalyst prepared from the HAlSiβ pristine support (Co<sub>1.8</sub>-HAlSiβ, **Figure 6c**) appeared to be the least active sample among the samples containing about 2 wt % Co. By comparing the initial C<sub>3</sub>H<sub>8</sub> conversion at 400 °C with that recorded at 400 °C after reaction at 500 °C, Co<sub>1.9</sub>-Siβ-pH-2.5 (**Figure 6a**) and Co<sub>1.8</sub>-HAlSiβ (**Figure 6c**) appeared to have suffered from deactivation after reaction at 500 °C by about 22 and 25 %, respectively (**Table 3**). Deactivation of these samples could be observed from 400 to 475 °C and 475 to 500 °C for Co<sub>1.9</sub>-Siβ-pH-2.5 (**Figure 6a**) and Co<sub>1.8</sub>-HAlSiβ (**Figure 6c**), respectively. In contrast, Co<sub>2.2</sub>-Siβ-pH-9.0 was found to be activated (**Figure 6b**) by about 13 % after reaction at 500 °C (**Table 3**). Except Co<sub>1.1</sub>-Siβ-pH-2.5, which was found to be stable, the other samples of the Co<sub>x</sub>-Siβ-pH-2.5 series were found to deactivate after reaction at 500 °C (**Table 3**). C<sub>3</sub>H<sub>8</sub> conversion at 400 °C after reaction at 500 °C was found to be close to 3 % with a selectivity in C<sub>3</sub>H<sub>6</sub> of about 50 % for a Co content lower than or equal to

3.0 wt % in the  $\text{Co}_x\text{-Si}\beta\text{-pH-2.5}$  series. Above this Co loading, the conversion of  $\text{C}_3\text{H}_8$  was found to decrease substantially and the selectivity in  $\text{C}_3\text{H}_6$  was found to decrease at the highest Co loading (9.3 wt %) only. Similar trends were also observed for the  $\text{Co}_x\text{-Si}\beta\text{-pH-9.0}$  series (**Table 3**), except that  $\text{C}_3\text{H}_8$  conversion and  $\text{C}_3\text{H}_6$  selectivity appeared to remain essentially constant up to a higher Co loading (4.3 wt %) compared to the  $\text{Co}_x\text{-Si}\beta\text{-pH-2.5}$  series (3.0 wt %). It can be noticed that  $\text{C}_3\text{H}_6$  selectivity increased substantially for catalysts exhibiting a deactivation greater than or equal to 16 % in the  $\text{Co}_x\text{-Si}\beta$  series (**Table 3**). This increase in  $\text{C}_3\text{H}_6$  selectivity cannot be attributed to a significant increase in the production of  $\text{C}_3\text{H}_6$  (**Table 4**). A close inspection of the evolution of the  $\text{CO}_x$  concentration with time on stream for  $\text{Co}_{0.9}\text{-Si}\beta\text{-pH-2.5}$  (**Figure 6a**) shows that the increase in  $\text{C}_3\text{H}_6$  selectivity from 36 to 47 % (**Table 3**) was mainly due to a significant decrease in the production of  $\text{CO}_x$  after reaction at 500 °C, as the concentration of  $\text{C}_3\text{H}_6$  was found to remain essentially constant (**Figure 6a, Table 3**). The sintering of part of the Co species could be ruled out as the origin of the decrease in  $\text{CO}_x$  production after ODHP reaction at 500 °C since the XPS Co/Si at. ratios of  $\text{Co}_{3.0}\text{Si}\beta\text{-pH-2.5}$  (Co/Si = 0.0099) and  $\text{Co}_{3.1}\text{Si}\beta\text{-pH-9.0}$  (Co/Si = 0.0118) were found to remain essentially constant after ODHP reaction (0.0107 and 0.0122, respectively). It must also be noticed that increasing amounts of  $\text{H}_2$ , most probably produced via the water-gas shift reaction [75], were recorded with increasing temperatures on all of the samples investigated in the present study, as also reported in earlier studies ([2] and references therein). In the present study, the production of  $\text{H}_2$  could be recorded even if complete conversion of  $\text{O}_2$  was not always achieved. In contrast, the production of  $\text{H}_2$  on V-Si $\beta$  catalysts may be attributed to the occurrence of the PDH reaction as  $\text{H}_2$  could only be observed at temperatures for which full  $\text{O}_2$  conversion was achieved [36]. More recently, some of us also recorded the production of  $\text{H}_2$  in the ODHP reaction over V-modified hydroxyapatite catalysts at temperatures for which full conversion of  $\text{O}_2$  was not achieved

[76]. In this latter study, the occurrence of the PDH reaction could be discarded since the removal of O<sub>2</sub> from the reacting feed led to almost complete disappearance of C<sub>3</sub>H<sub>6</sub> in the effluent, as is also the case over the Co<sub>x</sub>-Siβ materials prepared in the present study (**Figure S6**). The concentration of H<sub>2</sub> at 500 °C was found to be about 0.7 % for the Co<sub>x</sub>-Siβ-pH-2.5 samples containing from 1.1 to 5.6 wt % of Co, whereas Co<sub>9.3</sub>-Siβ-pH-2.5 produced 0.06 % of H<sub>2</sub> only. The amounts of H<sub>2</sub> produced on the Co<sub>x</sub>-Siβ-pH-9.0 series were found to be slightly higher than those found on the Co<sub>x</sub>-Siβ-pH-2.5 series, whereas that produced on Co<sub>1.8</sub>-HAlSiβ was much lower (0.2 %). These data show that the samples were exposed to hydrogen at relatively high temperatures in the course of the ODHP reaction. The significant increase in C<sub>3</sub>H<sub>6</sub> selectivity observed for samples exhibiting a deactivation greater than or equal to 16 % (**Table 3**), which is essentially attributed to a significant decrease in CO<sub>x</sub> production after reaction at 500 °C (**Figure 6a**), may be attributed to changes in the oxidation state and/or coking of some Co species (probably present in small quantities) under the experimental conditions of the ODHP reaction, such as CoO<sub>x</sub> species that have been reported to be fairly active in the oxidation of propane [77].

After reaction at 500 °C, the ODHP catalytic activity of Co<sub>1.9</sub>-Siβ-pH-2.5 was recorded at 400, 375 and 350 °C before being measured once more at 400 °C (**Figure S7**). The fact that the catalytic activity recorded at 400 °C after measurements at 375 and 350 °C was found to be identical to that recorded at 400 °C after reaction at 500 °C indicates that the measurements were performed on a stabilized catalytic sample. The apparent activation energies for the production of C<sub>3</sub>H<sub>6</sub> and CO<sub>x</sub> were estimated to be about 87 and 108 kJ mol<sup>-1</sup>, respectively, from the corresponding Arrhenius plots in the 350-400 °C temperature domain (**Figure S8**). The fact that the apparent activation energy for the production of C<sub>3</sub>H<sub>6</sub> is consistent with earlier data for this reaction [36] suggests that the rate measurements were performed in a kinetic regime free of diffusion limitation.



The turnover frequencies (TOFs), normalized by the total number of Co atoms (Section 2.3.) and determined at 400 °C after reaction at 500 °C, are listed in **Table 5**. This table shows firstly that Co<sub>1.8</sub>-HAISiβ was the least active sample ( $0.86 \times 10^{-3} \text{ s}^{-1}$ ) of the investigated catalysts. Moreover, it can be seen that the TOFs remained essentially constant and close to  $3.07 \pm 0.45 \times 10^{-3} \text{ s}^{-1}$  for samples with Co loadings below 5 wt % for both Co<sub>x</sub>-Siβ series, and were found to decrease at Co loadings higher than 3.0 and 4.3 wt% in the Co<sub>x</sub>-Siβ-pH-2.5 and Co<sub>x</sub>-Siβ-pH-9.0 series, respectively. It must be highlighted that the TOF measured in the present work over Co ( $3.07 \times 10^{-3} \text{ s}^{-1}$ ) is found to be about 7 times higher than that reported earlier for vanadium in comparable materials ( $0.44 \times 10^{-3} \text{ s}^{-1}$  [36]). The superiority of Co over V for ODHP is likely responsible for the interest in Co for this reaction [17-30]. The evolution of the TOF measured at 400 °C after ODHP reaction at 500 °C over the Co<sub>x</sub>-Siβ series (TOF essentially constant below 5 wt% Co and decreasing at higher Co contents, **Table 5**) cannot be correlated straightforwardly with the Co speciation deduced from H<sub>2</sub>-TPR (Section 3.2.3., **Figure 4**). This may be attributed to the fact that H<sub>2</sub>-TPR profiles were recorded on samples in their calcined state, whereas TOFs were estimated at 400 °C after ODHP reaction at 500 °C, which may have affected the Co speciation in the spent catalysts compared to that in the calcined samples. This may also be attributed to the fact that only a limited number of Co species were active in the ODHP reaction, as reported earlier in propane-SCR [71]. More interesting is that the above-mentioned evolution of the TOF in the Co<sub>x</sub>-Siβ samples can be correlated to the changes in C<sub>3</sub>H<sub>6</sub> production observed at 400 °C before and after ODHP reaction at 500 °C (**Table 4**). **Table 4** shows a slight increase in C<sub>3</sub>H<sub>6</sub> production after ODHP reaction at 500 °C for the Co<sub>x</sub>-Siβ samples with a Co loading below 5 wt%, whereas a decrease in C<sub>3</sub>H<sub>6</sub> production can be observed for Co loadings greater than 5 wt% indeed. The fact that the TOFs were found to remain essentially constant below 5 wt % of Co in the Co<sub>x</sub>-Siβ series

may be attributed to the maximum content of Co incorporable in the silanol nests created by dealumination of the HAlSi $\beta$  pristine support. On a one-to-one  $\text{Co}_{\text{inserted-to-Al}_{\text{extracted}}}$  basis, the decrease in Al content from 2.65 to 0.11 wt% after dealumination of the HAlSi $\beta$  support to Si $\beta$  (**Table 1**) would lead to a maximum Co loading of about 5.6 wt% indeed. Below the maximum content of Co incorporable in the silanol nests on both  $\text{Co}_x\text{-Si}\beta$  series, it may appear surprising that the  $\text{C}_3\text{H}_6$  TOF did not depend on the speciation of Co in the impregnating solutions (pH 2.5 and 9.0). **Figure 2** shows that Co speciation in the aqueous medium remained as  $\text{Co}(\text{H}_2\text{O})_6^{2+}$  at pH 2.5, whereas the precipitation of  $\text{Co}(\text{OH})_2$  was deduced to occur at pH 9.0, indeed. Such an unexpected result may be attributed to the transient decrease to pH 3.0 during the addition of the Si $\beta$  support to the Co solution at pH 9 (Section 2.1.). Under such acidic conditions, the  $\text{Co}(\text{OH})_2$  precipitate produced at pH 9.0 (**Figure 2**) may be dissolved rapidly to  $\text{Co}(\text{H}_2\text{O})_6^{2+}$ , allowing for the incorporation of Co in the Si $\beta$  framework as is the case in the preparation of the  $\text{Co}_x\text{-Si}\beta$  materials at pH 2.5. As the pH of the suspensions was subsequently adjusted back to pH 9.0 (Section 2.1.), the incorporation of Co in the framework of the Si $\beta$  support is assumed to be a rather rapid process, as reported for V incorporation in Si $\beta$ , which was shown to be kinetically promoted under acidic conditions [78].

**Table 6** lists the  $\text{C}_3\text{H}_6$  turnover frequencies (TOFs) on a per atom (Co or V) basis reported in earlier ODHP works at various temperatures. These earlier data are essentially compared to those estimated in the present study over the  $\text{Co}_{0.9}\text{-Si}\beta\text{-pH-2.5}$  sample, which was shown to be representative of the most active samples (**Table 5**). Firstly it can be noted that the TOF value at 400 °C reported by Bulánek et al. ( $0.75 \times 10^{-3} \text{ s}^{-1}$ , **Table 6** [19]) compares remarkably well with that obtained for our  $\text{Co}_{0.8}\text{-HAlSi}\beta$  sample (**Table 5**,  $0.86 \times 10^{-3} \text{ s}^{-1}$ ). These samples were (i) evaluated in the ODHP reaction under close experimental conditions, (ii) both synthesized by

cationic exchange of the protons of beta zeolites exhibiting close Si/Al ratios (17.6 for Bulánek et al. [19] vs 13.6 in the present study, **Table 1**) and (iii) exhibited close Co loadings (1.68 wt% Co for Bulánek et al. [19] vs 1.82 wt% Co in the present study, **Table 1**). This remarkable agreement provides support for the reliability of the rates measured in the ODHP reaction in the present study. Overall, **Table 6** shows that the TOFs of the Co<sub>1.9</sub>-Siβ-pH-2.5 sample are found to be greater than or equal to those reported in most of the earlier studies performed with Co-containing materials [17, 18, 20-28]. The fact that the TOF of Co<sub>1.9</sub>-Siβ-pH-2.5 appears to be lower than those reported on Co supported on Zr molecular organic framework (NU1000 [29]) and carbon nanotubes (CNTs [30]), respectively, illustrates the crucial importance of the support in the ODHP reaction. Regarding the earlier studies performed on silica-based materials, the TOF of the Co<sub>1.9</sub>-Siβ-pH-2.5 sample is found to be much higher than those reported in earlier studies [19, 23, 30] suggesting that the incorporation of the Co species into the framework of Siβ zeolite would be beneficial for the ODHP reaction.

#### 4. Conclusion

The Co<sub>x</sub>-Siβ (*x* being the Co weight loading) samples prepared by incorporation of Co<sup>2+</sup> species in the silanol nests (created by dealumination of a HAlSiβ support) appeared to be about 4 times more active in the ODHP reaction on a TOF basis compared to a reference sample for which Co was located in cationic exchange positions of a HAlSiβ support. Despite obvious differences in Co speciation in solution between pH 2.5 and 9.0 (**Figure 2**), the change in pH during deposition of the Co<sup>2+</sup> species in the silanol nests was found to have hardly any influence on the resulting ODHP catalytic activity below a Co loading of 5 wt%. This peculiarity was assigned to the transient

acidification of the medium to pH 3.0 with the addition of Si $\beta$  in the Co solution set originally at pH 9.0, and the rapid dissolution of the Co(OH)<sub>2</sub> precipitate and incorporation of the Co species in the Si $\beta$  framework. The fact that the TOF was found to remain essentially constant below 5 wt % of Co in the Co<sub>x</sub>-Si $\beta$  series may be rather attributed to the maximum content of Co incorporable in the silanol nests created by dealumination of the HAlSi $\beta$  pristine support. The TOF of a Co<sub>1.9</sub>-Si $\beta$ -pH-2.5 sample is found to be greater than or equal to those reported in most of the earlier studies performed with Co-containing materials and to be about 7 times greater than that of V-Si $\beta$  materials [36] under comparable experimental conditions in terms of O<sub>2</sub> and C<sub>3</sub>H<sub>8</sub> concentrations in the reacting feed (**Table 6**). The latter illustrates the superiority of Co over V for the ODHP reaction when incorporated into the silanol nests of a dealuminated Si $\beta$  supporting oxide. Finally, this work paves the way for the characterization of the Co<sup>2+</sup> species by NO-TPD. NO was found to be weakly chemisorbed on the Co<sup>2+</sup> species and to be relatively sensitive to the Co environment as the NO/Co ratio was found to decrease as the loading of Co increased, however, whereas differences in Co dispersion could not be observed by XPS. Additional work is currently under progress on more model materials for the characterization of supported Co<sup>2+</sup> species by NO-TPD.

#### **CRedit author statement**

**Stanislaw Dzwigaj:** Investigation, Resources, Writing – Review & Editing. **Diane Reja:** Investigation, Data Curation. **Saremlé Koné-Guira:** Investigation, Data Curation. **Antoine Miche:** Investigation, Data Curation. **Guylène Costentin:** Conceptualization, Writing – Review & Editing. **Cyril Thomas:** Conceptualization, Validation, Investigation, Resources, Data Curation,

Writing – Original Draft, Writing – Review & Editing, Visualization, Supervision, Project Administration. All co-authors have read and agreed on the submitted version of the manuscript.

### **Acknowledgments**

The authors thank Sorbonne Université and CNRS for providing funding (Crédits récurrents) and access to the analytical equipment used in the present work. D.R. gratefully acknowledges Sorbonne Université for financial support (PhD Grant no 3892/2020). The authors also express their gratitude to J. Noël (Plateforme ALIPP6, Laboratoire d'Analyse Géochimie Élémentaire, Sorbonne Université) for her rigorous analysis of the ICP-OES data. Dr. C. Reynaud (Sorbonne Université, PhD Grant no 3415/2019) is acknowledged for the computation of the data regarding Co speciation in solution shown in **Figure 2**. The authors also acknowledge L. Anscieau for his help in recovering the original data of **Figure S5** that were stored in the hard disk drive of an out of order 27 years old PC.

### **Conflicts of interest**

There is no conflict of interest to declare.

**Table 1.** Chemical composition of the samples determined by ICP-OES.

Samples	Targeted Co	Element concentration			Si/Al at. ratio
	loading (wt%)	Si (wt%)	Al (wt%)	Co (wt%)	
HAISi $\beta$	-	37.63	2.65	-	13.6
Co <sub>1.8</sub> -HAISi $\beta$	2.0	37.97	2.63	1.82	13.9
Si $\beta$	-	42.00	0.11	-	380
Co <sub>1.1</sub> -Si $\beta$ -pH-2.5	1.0	41.47	0.09	1.11	459
Co <sub>1.9</sub> -Si $\beta$ -pH-2.5	2.0	43.02	0.09	1.89	470
Co <sub>3.0</sub> -Si $\beta$ -pH-2.5	3.0	42.03	0.08	2.95	508
Co <sub>5.6</sub> -Si $\beta$ -pH-2.5	5.0	40.92	0.08	5.64	501
Co <sub>9.3</sub> -Si $\beta$ -pH-2.5	9.0	39.25	0.08	9.26	467
Co <sub>0.9</sub> -Si $\beta$ -pH-9.0	1.0	44.15	0.07	0.93	579
Co <sub>2.2</sub> -Si $\beta$ -pH-9.0	2.0	43.12	0.09	2.19	488
Co <sub>3.1</sub> -Si $\beta$ -pH-9.0	3.0	42.32	0.09	3.07	469
Co <sub>4.3</sub> -Si $\beta$ -pH-9.0	5.0	41.97	0.07	4.34	565
Co <sub>9.3</sub> -Si $\beta$ -pH-9.0	9.0	39.43	0.08	9.28	487

**Table 2.** O 1s, C 1s, Al 2p and Co 2p<sub>3/2</sub> binding energies referenced to Si 2p at 103.3 eV and Si/Al at. ratio.

Samples	Binding energy (eV)				Si/Al at. ratio
	O 1s	C 1s	Al 2p	Co 2p <sub>3/2</sub>	
HAISiβ	532.6	284.4	74.5	-	15.3
Co <sub>1.8</sub> -HAISiβ	532.5	284.9	74.5	782.5	15.6
Siβ	532.6	283.9	*	-	-
Co <sub>1.1</sub> -Siβ-pH-2.5	532.5	283.8	*	781.4	-
Co <sub>1.9</sub> -Siβ-pH-2.5	532.5	283.8	*	781.2	-
Co <sub>3.0</sub> -Siβ-pH-2.5	532.5	284.0	*	781.4	-
Co <sub>5.6</sub> -Siβ-pH-2.5	532.5	284.3	*	781.3	-
Co <sub>9.3</sub> -Siβ-pH-2.5	532.5	283.9	*	781.3	-
Co <sub>0.9</sub> -Siβ-pH-9.0	532.5	284.0	*	781.3	-
Co <sub>2.2</sub> -Siβ-pH-9.0	532.5	283.8	*	781.2	-
Co <sub>3.1</sub> -Siβ-pH-9.0	532.6	285.0	*	781.8	-
Co <sub>4.3</sub> -Siβ-pH-9.0	532.5	283.9	*	780.9	-
Co <sub>9.3</sub> -Siβ-pH-9.0	532.5	283.9	*	780.8	-

\*Absence of Al 2p signal measurable under the conditions used.

**Table 3.** Initial C<sub>3</sub>H<sub>8</sub> conversion (X<sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup>) at 400 °C and C<sub>3</sub>H<sub>8</sub> conversion at 400 °C after reaction at 500 °C (X<sub>C<sub>3</sub>H<sub>8</sub><sup>f</sup>) together with the corresponding selectivity in C<sub>3</sub>H<sub>6</sub> (S<sub>C<sub>3</sub>H<sub>6</sub><sup>i</sup> and S<sub>C<sub>3</sub>H<sub>6</sub><sup>f</sup>) in the C<sub>3</sub>H<sub>8</sub> ODHP reaction (8.2 % C<sub>3</sub>H<sub>8</sub> - 1.9 % O<sub>2</sub> - He, about 20 μmol Co in the sample weight (sw) of catalyst diluted in 600 mg α-Al<sub>2</sub>O<sub>3</sub>, 74.1 mL min<sup>-1</sup> total flow rate) over various Co samples. X<sub>C<sub>3</sub>H<sub>8</sub><sup>diff</sup> represents the evolution of the C<sub>3</sub>H<sub>8</sub> conversion at 400 °C after reaction at 500 °C (X<sub>C<sub>3</sub>H<sub>8</sub><sup>diff</sup> = ((X<sub>C<sub>3</sub>H<sub>8</sub><sup>f</sup> - X<sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup>) / X<sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup>) × 100).</sub></sub></sub></sub></sub></sub></sub></sub></sub>

sw (mg)	x wt %	Co <sub>x</sub> -Siβ-pH-2.5					Co <sub>x</sub> -HAISiβ					Co <sub>x</sub> -Siβ-pH-9.0						
		X <sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>i</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>f</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>f</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>diff</sup> %</sub>	x wt %	X <sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>i</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>f</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>f</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>diff</sup> %</sub>	x wt %	X <sub>C<sub>3</sub>H<sub>8</sub><sup>i</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>i</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>f</sup> %</sub>	S <sub>C<sub>3</sub>H<sub>6</sub><sup>f</sup> %</sub>	X <sub>C<sub>3</sub>H<sub>8</sub><sup>diff</sup> %</sub>
120	1.1	2.8	47	2.8	48	0							0.9	3.0	45	3.2	46	+ 6
60	1.9	3.7	36	2.9	47	- 22	1.8	0.8	67	0.6	60	- 25	2.2	2.4	47	2.7	49	+ 13
40	3.0	3.7	34	2.9	45	- 22							3.1	3.1	44	2.8	49	- 11
24	5.6	2.2	53	2.0	55	- 10							4.3	3.8	36	3.2	44	- 16
13	9.3	1.9	33	1.6	38	- 16							9.3	2.8	33	2.0	38	- 30



**Table 4.** Initial C<sub>3</sub>H<sub>6</sub> concentration at 400 °C (C<sub>3</sub>H<sub>6</sub><sup>i</sup>) and C<sub>3</sub>H<sub>6</sub> concentration at 400 °C after ODHP reaction at 500 °C (C<sub>3</sub>H<sub>6</sub><sup>f</sup>) (8.2 % C<sub>3</sub>H<sub>8</sub> - 1.9 % O<sub>2</sub> - He, about 20 μmol Co diluted in 600 mg α-Al<sub>2</sub>O<sub>3</sub>, 74.1 mL min<sup>-1</sup> total flow rate).

Co <sub>x</sub> -Siβ-pH-2.5			Co <sub>x</sub> -HAISiβ			Co <sub>x</sub> -Siβ-pH-9.0		
<i>x</i> wt%	C <sub>3</sub> H <sub>6</sub> <sup>i</sup> (ppm)	C <sub>3</sub> H <sub>6</sub> <sup>f</sup> (ppm)	<i>x</i> wt%	C <sub>3</sub> H <sub>6</sub> <sup>i</sup> (ppm)	C <sub>3</sub> H <sub>6</sub> <sup>f</sup> (ppm)	<i>x</i> wt%	C <sub>3</sub> H <sub>6</sub> <sup>i</sup> (ppm)	C <sub>3</sub> H <sub>6</sub> <sup>f</sup> (ppm)
1.1	1051	1085	1.8	570	501	0.9	976	1134
1.9	1079	1093				2.2	928	1067
3.0	1041	1105				3.1	900	1071
5.6	932	881				4.3	1095	1128
9.3	501	464				9.3	751	610

**Table 5.** Turnover frequency (TOF [54]) data calculated on the basis of the total number of Co atoms and on the amount of C<sub>3</sub>H<sub>6</sub> produced at 400 °C after ODHP reaction at 500 °C (8.2 % C<sub>3</sub>H<sub>8</sub> - 1.9 % O<sub>2</sub> - He, about 20 μmol Co diluted in 600 mg α-Al<sub>2</sub>O<sub>3</sub>, 74.1 mL min<sup>-1</sup> total flow rate).

TOF (x 10 <sup>3</sup> s <sup>-1</sup> )		TOF (x 10 <sup>3</sup> s <sup>-1</sup> )		TOF (x 10 <sup>3</sup> s <sup>-1</sup> )	
<i>x</i>	Co <sub><i>x</i></sub> -Siβ-pH-2.5	<i>x</i>	Co <sub><i>x</i></sub> -HAlSiβ	<i>x</i>	Co <sub><i>x</i></sub> -Siβ-pH-9.0
1.1	2.65			0.9	3.47
1.9	3.13	1.8	0.86	2.2	2.64
3.0	3.04			3.1	3.01
5.6	2.11			4.3	3.52
9.3	1.21			9.3	1.60

**Table 6.** Comparison of the turnover frequencies (TOFs, molecule C<sub>3</sub>H<sub>6</sub> s<sup>-1</sup> Co or V atom<sup>-1</sup>) reported in earlier studies with that measured in the present work on Co<sub>1.9</sub>-Siβ-pH-2.5. These TOF calculated on a per atom basis were either extracted as is from the corresponding studies<sup>i</sup> or estimated from the data reported in these earlier works<sup>j</sup>.

Sample	Co or V (wt%)	sw <sup>a</sup> (mg)	[C <sub>3</sub> H <sub>8</sub> ] <sup>b</sup> (%)	[O <sub>2</sub> ] <sup>c</sup> (%)	F <sup>d</sup> (mL/min)	T <sup>e</sup> (°C)	X C <sub>3</sub> H <sub>8</sub> <sup>f</sup> (%)	S C <sub>3</sub> H <sub>6</sub> <sup>g</sup> (%)	TOF <sup>h</sup> (x 10 <sup>3</sup> s <sup>-1</sup> )	Refs.
Co-AIM + NU1000	-	80-100	3.0	10.0	-	230	-	-	0.28 <sup>i</sup>	[29]
Co/ZrO <sub>2</sub>	-	-	3.0	10.0	-	230	-	-	0.04 <sup>i</sup>	[29]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>230</b>	-	-	0.02 <sup>k</sup>	This work
CeCo <sub>0.5</sub> O <sub>y</sub>	9.72	100	5.0	15.0	100	300	-	-	0.25 <sup>j</sup>	[18]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>300</b>	-	-	0.21 <sup>k</sup>	This work
Cs <sub>2.5</sub> H <sub>6x-0.5</sub> Co <sub>1-x</sub> PV <sub>1</sub> Co <sub>x</sub> Mo <sub>11-x</sub> O <sub>40</sub>	2.81	500-750	40.0	20.0	15	345	5.0	79.1	0.74-0.49 <sup>j</sup>	[20]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>345</b>	-	-	0.79 <sup>k</sup>	This work
CoOx-CNTs	0.49	-	16.7	16.7	30	353	-	71	94.55 <sup>j</sup>	[30]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>353</b>	-	-	0.98 <sup>k</sup>	This work
CoOx-MCM-41	0.49	-	16.7	16.7	30	393	-	-	0.70 <sup>j</sup>	[30]
CoOx-SiO <sub>2</sub>	0.49	-	16.7	16.7	30	393	-	-	0.33 <sup>j</sup>	[30]
CoH-BEA-17.5-0.42	1.68	200	5.0	6.5	100	400	3.9	29.4	0.75 <sup>k</sup>	[19]
CoMO $\phi$ <sub>y</sub> (CoMoO <sub>4</sub> )	26.9	200	29.0	14.9	20	400	8.8	46.4	0.19 <sup>j</sup>	[26]
Co <sup>III</sup> <sub>0.36</sub> Co <sup>II</sup> <sub>0.30</sub> Al <sub>0.34</sub> (PO <sub>4</sub> ) <sub>0.11</sub> O <sub>1.19</sub>	50.1	100	16.7	16.7	50	400	27.9	36.0	0.73 <sup>j</sup>	[28]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>400</b>	2.9	46.8	3.13	This work
V <sub>0.25</sub> -Siβ	0.25	200	6.6	2.1	15	400	1.3	34.1	0.4 <sup>i</sup>	[36]
Co36SrHAp	1.42	500	14.3	4.0	30	450	15.8	46.6	1.95 <sup>j</sup>	[21]

Co <sub>2</sub> O <sub>3</sub> /SiO <sub>2</sub>	3.55	1000	18.1	18.1	83	450	15.8	15.8	0.46 <sup>j</sup>	[23]
CoMgAlO	9.67	-	-	-	-	450	1.5	92.5	0.33 <sup>j</sup>	[24]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>450</b>	-	-	9.13 <sup>k</sup>	This work
CoWO <sub>4</sub>	19.2	1000	15.0	15.0	60	500	12.6	38.5	0.10 <sup>j</sup>	[17]
Co <sub>2</sub> B/Co <sub>3</sub> B	93.2	235	30.0	15.0	50-160	500	13-23	88.2	0.36 <sup>j</sup>	[27]
P(0.05)CoO	69.4	100	11.8	9.8	50	500	27.2	53.8	0.55 <sup>j</sup>	[25]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>500</b>	-	-	23.18 <sup>k</sup>	This work
CoMOφx (CoMoO <sub>4</sub> )	29.8	1000	29.0	14.9	100	600	15.0	38.5	0.25 <sup>j</sup>	[22]
Co <sub>1.9</sub> -Siβ-pH-2.5	1.89	60	8.2	1.9	74.1	<b>600</b>	-	-	108.50 <sup>k</sup>	This work

<sup>a</sup> sample weight, <sup>b</sup> C<sub>3</sub>H<sub>8</sub> concentration in the feed as vol. %, <sup>c</sup> O<sub>2</sub> concentration in the feed as vol. %, <sup>d</sup> even though in most studies the fact that the flow rates were not explicitly defined in the normal conditions (i.e. 273 K and 1 atm), the reported flow rates were considered as such for the estimation of the TOF, <sup>e</sup> reaction temperature, <sup>f</sup> conversion of C<sub>3</sub>H<sub>8</sub>, <sup>g</sup> Selectivity in C<sub>3</sub>H<sub>6</sub>, <sup>h</sup> TurnOver Frequency on a per cobalt atom basis, TOF extracted as is from the corresponding studies<sup>i</sup> or estimated from the data reported in these earlier works<sup>j</sup>, <sup>k</sup> TOF values extrapolated at the corresponding temperatures from the TOF estimated at 400 °C (3.13 x 10<sup>-3</sup> s<sup>-1</sup>) and considering an activation energy of 86.6 kJ/mol (Figure S8a).

### Figure Captions:

**Figure 1.** NO<sub>x</sub> profiles recorded by chemiluminescence analyzers (Thermo Environmental Instruments 42C-HL (red trace) or Eco Physics CLD 700 AL (black trace)) on about 0.12 g of Co<sub>1.1</sub>Siβ-pH-2.5 after annealing of the sample at 500 °C for 2 h in 100 mL<sub>NTP</sub>/min of He or Ar and RT exposure to 230 mL<sub>NTP</sub>/min of NO (400 ppm) in He or Ar for 2 h, followed by temperature-programmed desorption from RT to 400 °C (3 °C/min) in 230 mL<sub>NTP</sub>/min of He or Ar.

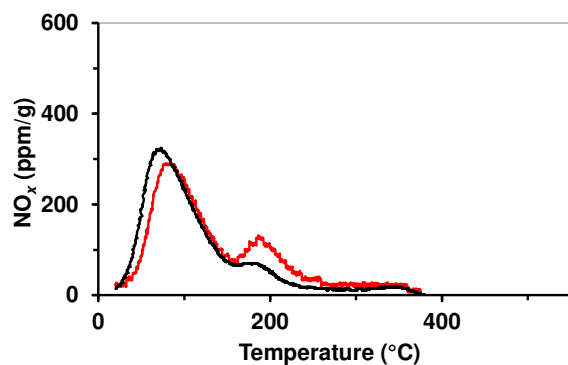
**Figure 2** Speciation of Co as a function of the pH of the aqueous medium for Co precursor concentrations of (a) 1.6 and (b) 17.4 × 10<sup>-3</sup> mol/L. Dotted lines put emphasis on the Co species present at pH 2.5 and 9.0.

**Figure 3.** Co 2p<sub>3/2</sub>/Si 2p surface at. (XPS) ratios as a function of the bulk (ICP-OES) ones. Co<sub>x</sub>-Siβ-pH-2.5 (○), Co<sub>x</sub>-Siβ-pH-9.0 (●) and Co<sub>1.8</sub>-HAlSiβ (●).

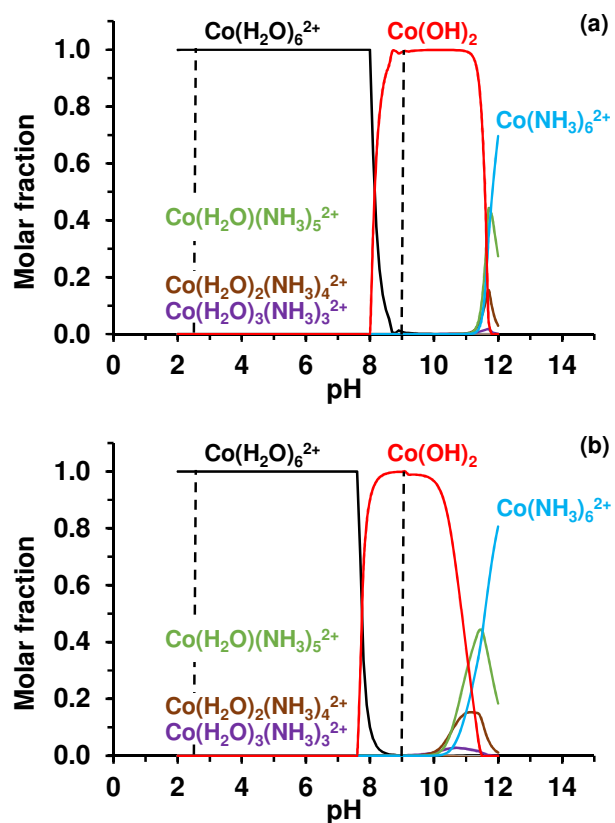
**Figure 4.** H<sub>2</sub>-TPR profiles recorded from RT to 900 °C in 5% H<sub>2</sub>/Ar with a 25 mL/min flow rate and a heating rate of 10 °C/min on (a) Co<sub>x</sub>-Siβ-pH-2.5 and (b) Co<sub>x</sub>-Siβ-pH-9.0 previously calcined at 500 °C (10 °C/min) under 5%O<sub>2</sub>/He for 1 h. It must be noted that the H<sub>2</sub>-TPR profile could not be recorded on Co<sub>9.3</sub>-Siβ-pH-2.5 because of an experimental issue in the course of the H<sub>2</sub>-TPR experiment on the last aliquot of the corresponding sample.

**Figure 5.** NO-TPD profiles recorded on (a) HAlSi $\beta$  and Co<sub>1.8</sub>-HAlSi $\beta$ , (b) Co<sub>x</sub>-Si $\beta$ -pH-2.5 and (c) Co<sub>x</sub>-Si $\beta$ -pH-9.0 after annealing of the samples at 500 °C for 2 h in 100 mL<sub>NTP</sub> min<sup>-1</sup> of He or Ar and RT exposure to 230 mL<sub>NTP</sub> min<sup>-1</sup> of NO(400 ppm) in He or Ar for 2 h followed by temperature-programmed desorption from RT to 500 °C (3 °C min<sup>-1</sup>) in 230 mL<sub>NTP</sub> min<sup>-1</sup> of He or Ar. (d) NO uptakes deduced from the corresponding NO-TPD profiles and (e) NO/Co ratios for Co<sub>x</sub>-Si $\beta$ -pH-2.5 and Co<sub>x</sub>-Si $\beta$ -pH-9.0 samples.

**Figure 6.** Evolution of the conversion of C<sub>3</sub>H<sub>8</sub> (black trace) and the concentrations in C<sub>3</sub>H<sub>6</sub> (blue trace), O<sub>2</sub> (red trace) and CO<sub>x</sub> (gray trace) (8.2% C<sub>3</sub>H<sub>8</sub>, 1.9% O<sub>2</sub> in He, 74.1 mL<sub>NTP</sub> min<sup>-1</sup> total flow) in the ODHP reaction as a function of time on stream and reacting temperature on 60 mg of catalyst diluted in 600 mg of  $\alpha$ -Al<sub>2</sub>O<sub>3</sub>, (a) Co<sub>1.9</sub>-Si $\beta$ -pH-2.5, (b) Co<sub>2.2</sub>-Si $\beta$ -pH-9.0 and (c) Co<sub>1.8</sub>-HAlSi $\beta$ . The numbers reported on the figures are related to the reaction temperature.

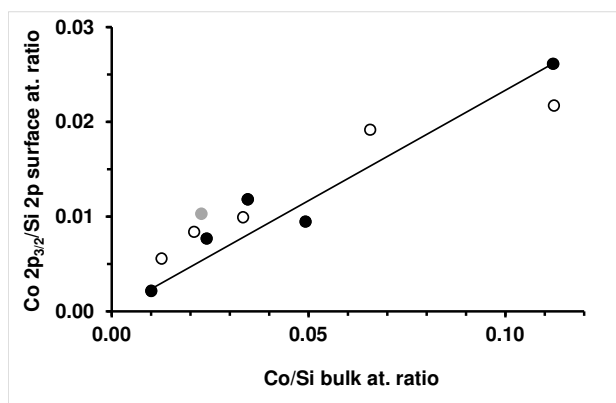


**Figure 1.** NO<sub>x</sub> profiles recorded by chemiluminescence analyzers (Thermo Environmental Instruments 42C-HL (red trace) or Eco Physics CLD 700 AL (black trace)) on about 0.12 g of Co<sub>1.1</sub>Siβ-pH-2.5 after annealing of the sample at 500 °C for 2 h in 100 mL<sub>NTP</sub>/min of He or Ar and RT exposure to 230 mL<sub>NTP</sub>/min of NO (400 ppm) in He or Ar for 2 h, followed by temperature-programmed desorption from RT to 400 °C (3 °C/min) in 230 mL<sub>NTP</sub>/min of He or Ar.

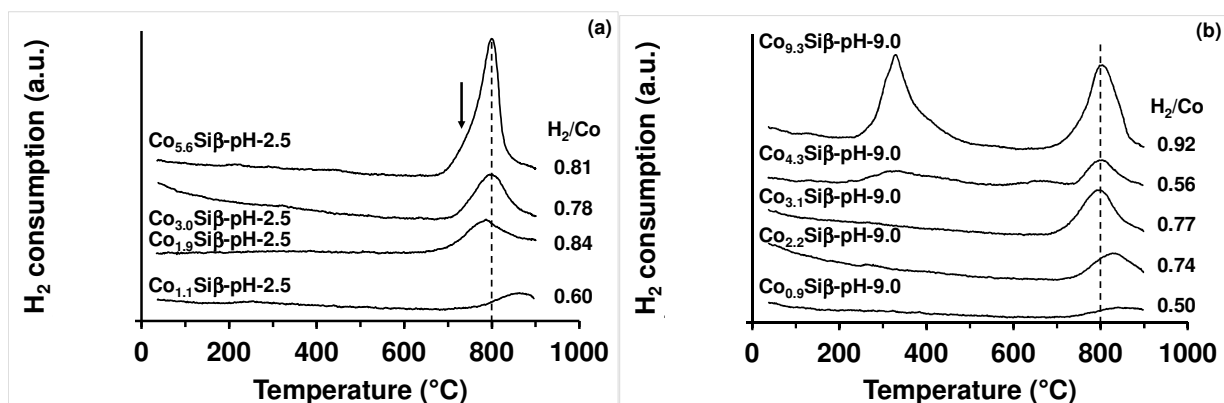


**Figure 2.** Speciation of Co as a function of the pH of the aqueous medium for Co precursor concentrations of (a)  $1.6$  and (b)  $17.4 \times 10^{-3}$  mol/L. Dotted lines put emphasis on the Co species present at pH 2.5 and 9.0.

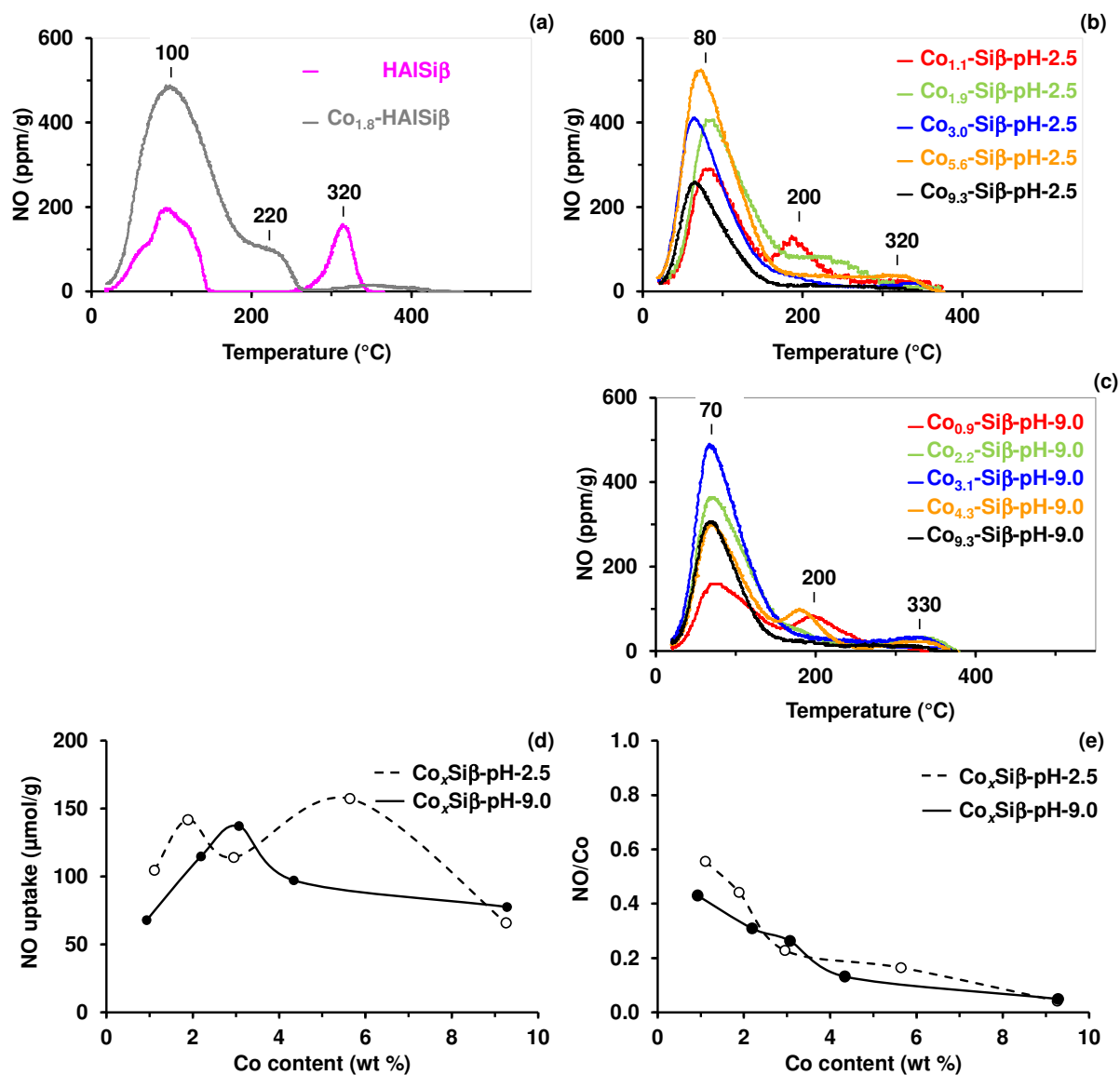




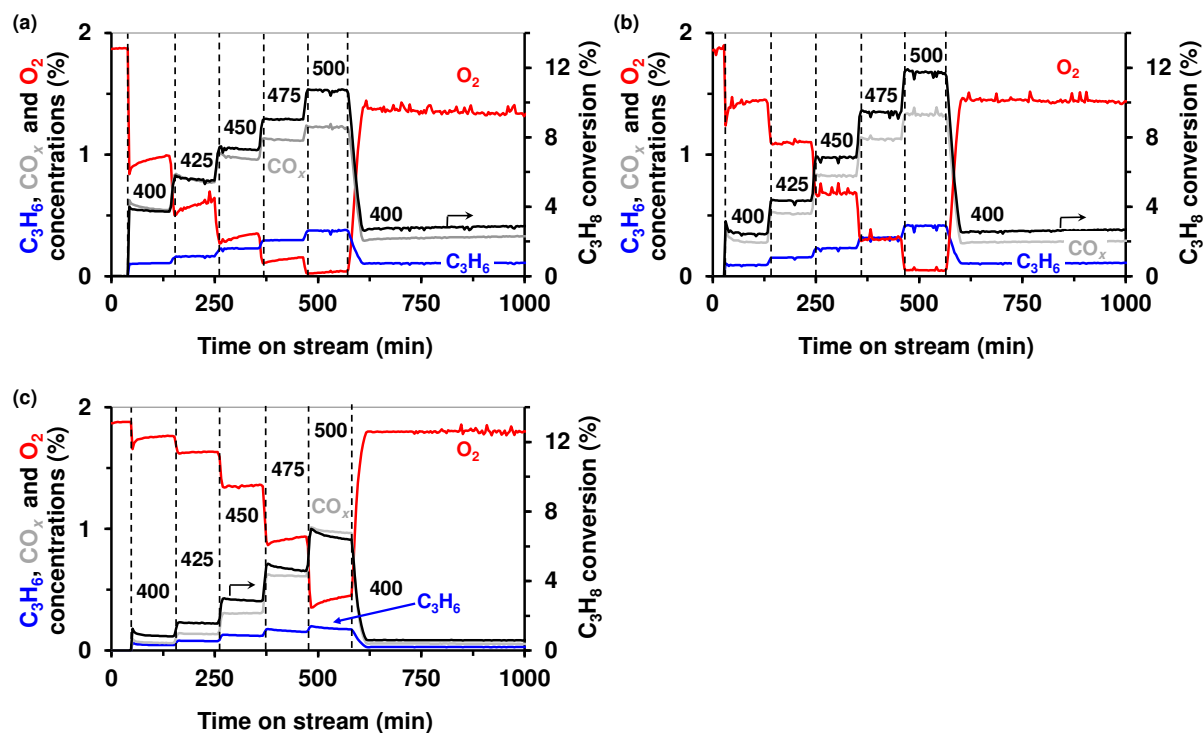
**Figure 3.** Co 2p<sub>3/2</sub>/Si 2p surface at. (XPS) ratios as a function of the bulk (ICP-OES) ones. Co<sub>x</sub>-Siβ-pH-2.5 (○), Co<sub>x</sub>-Siβ-pH-9.0 (●) and Co<sub>1.8</sub>-HAlSiβ (●).



**Figure 4.** H<sub>2</sub>-TPR profiles recorded from RT to 900 °C in 5% H<sub>2</sub>/Ar with a 25 mL/min flow rate and a heating rate of 10 °C/min on (a) Co<sub>x</sub>-Siβ-pH-2.5 and (b) Co<sub>x</sub>-Siβ-pH-9.0 previously calcined at 500 °C (10 °C/min) under 5%O<sub>2</sub>/He for 1 h. It must be noted that the H<sub>2</sub>-TPR profile could not be recorded on Co<sub>9.3</sub>-Siβ-pH-2.5 because of an experimental issue in the course of the H<sub>2</sub>-TPR experiment on the last aliquot of the corresponding sample.



**Figure 5.** NO-TPD profiles recorded on (a) HAISiβ and Co<sub>1.8</sub>-HAISiβ, (b) Co<sub>x</sub>-Siβ-pH-2.5 and (c) Co<sub>x</sub>-Siβ-pH-9.0 after annealing of the samples at 500 °C for 2 h in 100 mL<sub>NTP</sub> min<sup>-1</sup> of He or Ar and RT exposure to 230 mL<sub>NTP</sub> min<sup>-1</sup> of NO(400 ppm) in He or Ar for 2 h, followed by temperature-programmed desorption from RT to 500 °C (3 °C min<sup>-1</sup>) in 230 mL<sub>NTP</sub> min<sup>-1</sup> of He or Ar. (d) NO uptakes deduced from the corresponding NO-TPD profiles and (e) NO/Co ratios for Co<sub>x</sub>-Siβ-pH-2.5 and Co<sub>x</sub>-Siβ-pH-9.0 samples.



**Figure 6.** Evolution of the conversion of  $C_3H_8$  (black trace) and the concentrations in  $C_3H_6$  (blue trace),  $O_2$  (red trace) and  $CO_x$  (gray trace) (8.2%  $C_3H_8$ , 1.9%  $O_2$  in He, 74.1 mL<sub>NTP</sub> min<sup>-1</sup> total flow) in the ODHP reaction as a function of time on stream and reacting temperature on 60 mg of catalyst diluted in 600 mg of  $\alpha$ - $Al_2O_3$ , (a)  $Co_{1.9}$ -Si $\beta$ -pH-2.5, (b)  $Co_{2.2}$ -Si $\beta$ -pH-9.0 and (c)  $Co_{1.8}$ -HAISi $\beta$ . The numbers reported on the figures are related to the reaction temperature.

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# Supporting Information

## Cobalt on Dealuminated-Si $\beta$ as a Catalyst for the Oxidative Dehydrogenation of Propane

Stanislaw Dzwigaj, Diane Reja, Saremlé Koné-Guira, Antoine Miche,

Guylène Costentin and Cyril Thomas\*

Sorbonne Université, CNRS, Laboratoire Réactivité de Surface, LRS, F-75005 Paris, France.

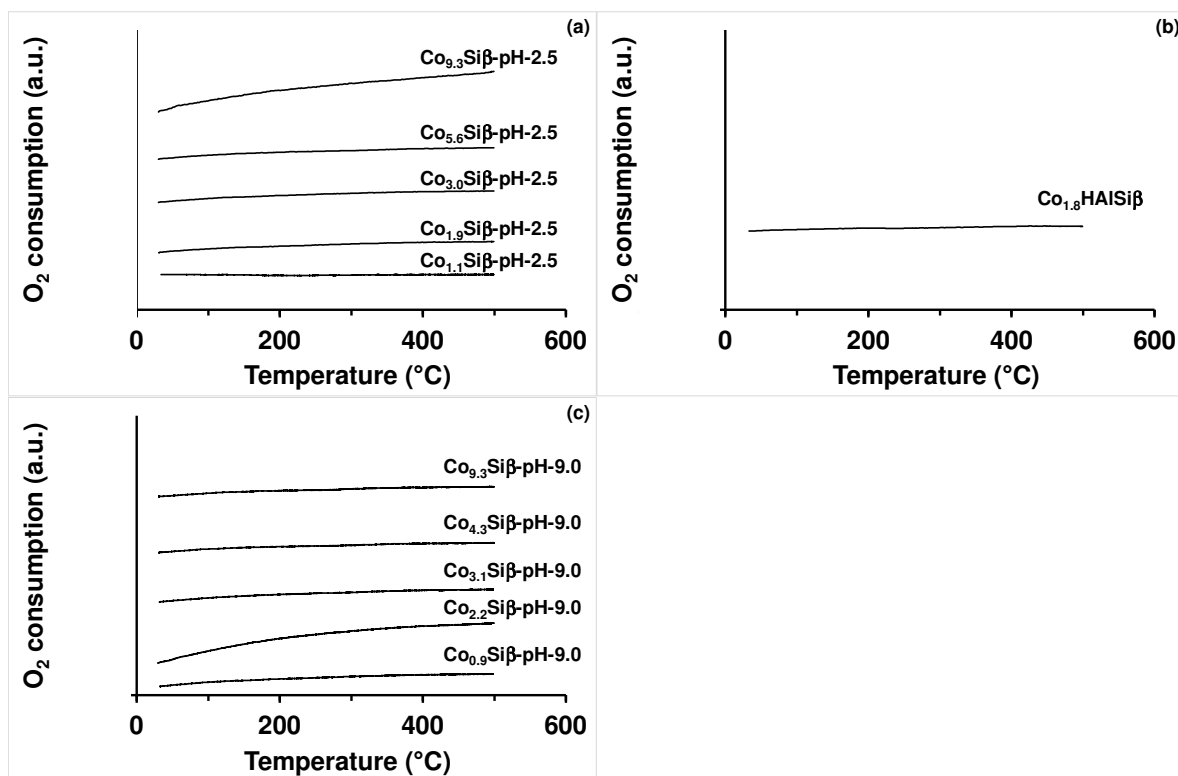
\* To whom correspondence should be addressed:

Dr. Cyril Thomas

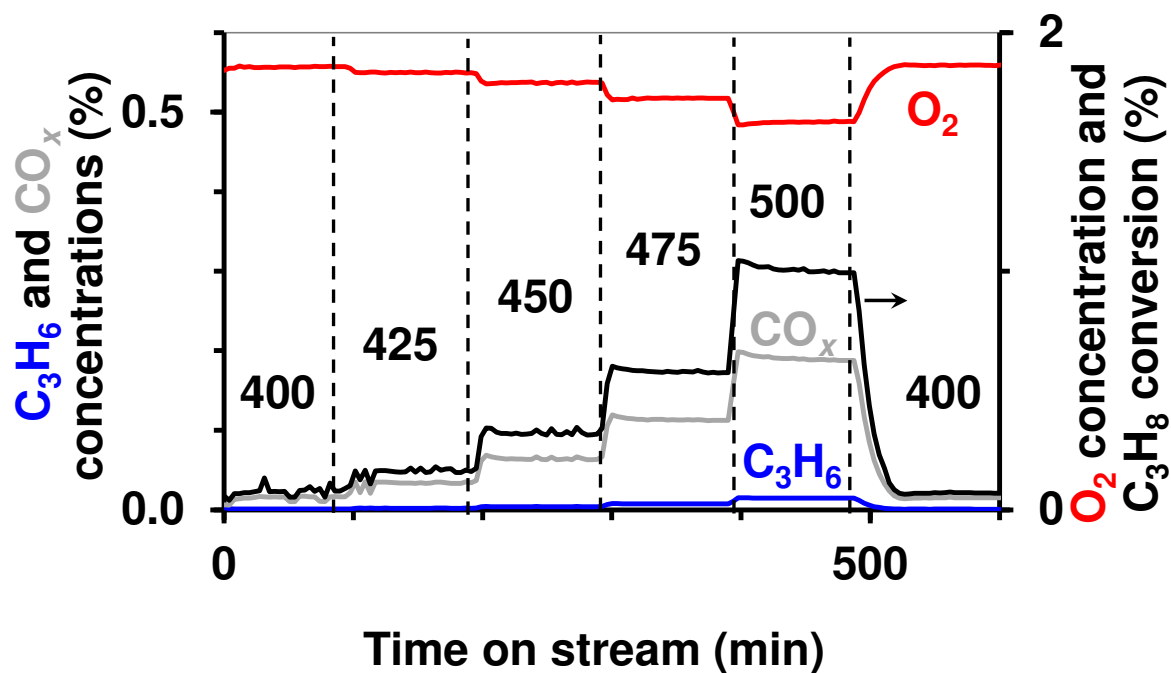
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Jussieu, Tour 43-53, 3<sup>ème</sup> étage, Case 178, F-75252, Paris, France

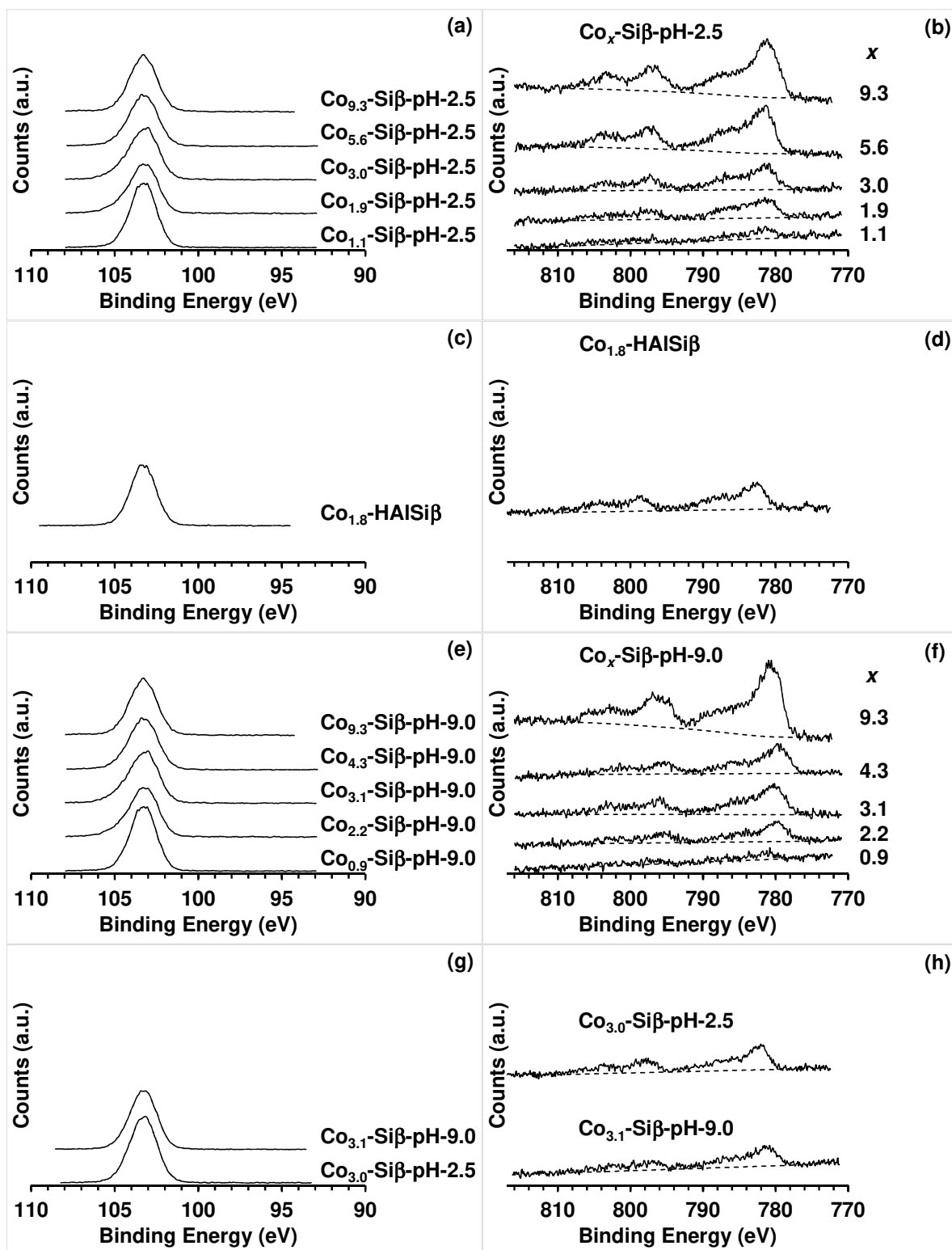
e-mail: [cyril.thomas@sorbonne-universite.fr](mailto:cyril.thomas@sorbonne-universite.fr)



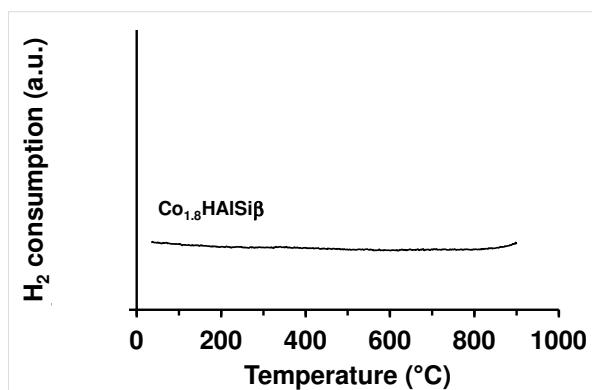
**Figure S1.** O<sub>2</sub>-TPO profiles recorded from RT to 500 °C in 5% O<sub>2</sub>/He with a 25 mL/min flow rate and a heating rate of 10 °C/min on (a) Co<sub>x</sub>-Siβ-pH-2.5, (b) Co<sub>1.8</sub>-HAlSiβ and (c) Co<sub>x</sub>-Siβ-pH-9.0.



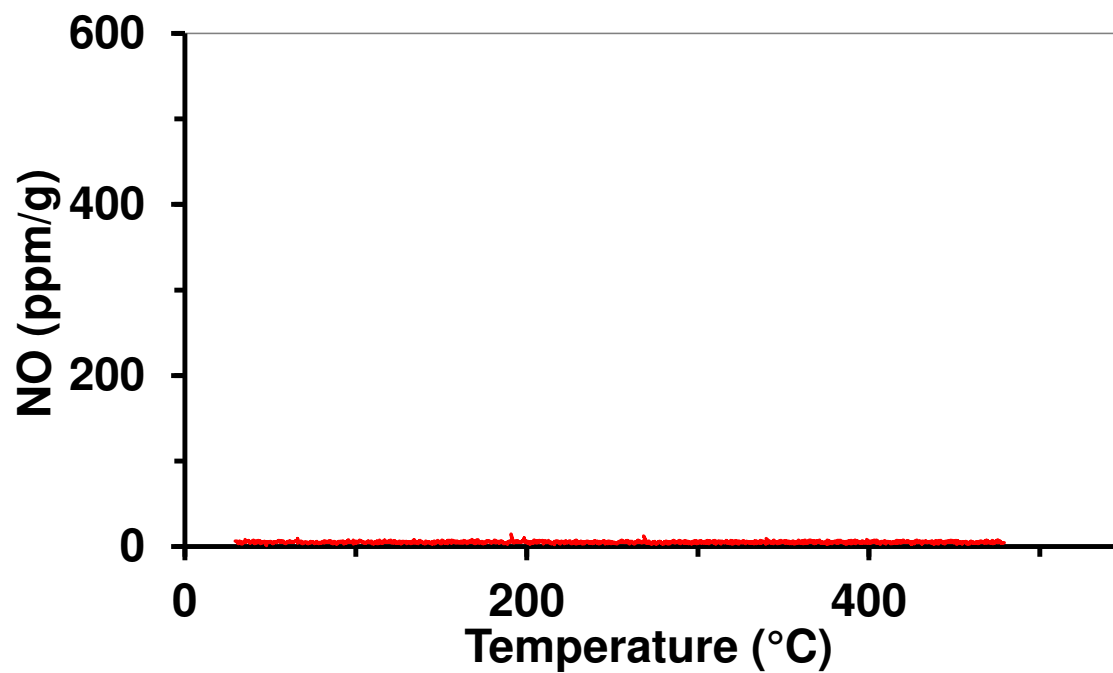
**Figure S2.** Evolution of the conversion of  $C_3H_8$  (black trace) and the concentrations in  $C_3H_6$  (blue trace),  $O_2$  (red trace) and  $CO_x$  (gray trace) (8.2%  $C_3H_8$ , 1.9%  $O_2$  in He,  $74.1 \text{ mL}_{NTP} \text{ min}^{-1}$  total flow) as a function of time on stream and reacting temperature on 600 mg of  $\alpha\text{-Al}_2\text{O}_3$ . The numbers reported on the figure are related to the reaction temperature.



**Figure S3.** Si 2p (a, c, e, g) and Co 2p (b, d, f, h) core levels of the various samples before (a-f) and after (g,h) ODHP reaction.



**Figure S4.** H<sub>2</sub>-TPR profile recorded from RT to 900 °C in 5% H<sub>2</sub>/Ar with a 25 mL/min flow rate and a heating rate of 10 °C/min on Co<sub>1.8</sub>-HAISiβ calcined at 500 °C (10 °C/min) under 5%O<sub>2</sub>/He for 1 h.



**Figure S5.** NO-TPD profile recorded on 120 mg of Si $\beta$  after annealing of the sample at 500 °C for 2 h in 100 mL<sub>NTP</sub> min<sup>-1</sup> of He and RT exposure to 230 mL<sub>NTP</sub> min<sup>-1</sup> of NO(400 ppm) in He for 2 h followed by temperature-programmed desorption from RT to 500 °C (3 °C min<sup>-1</sup>) in 230 mL<sub>NTP</sub> min<sup>-1</sup> of He.

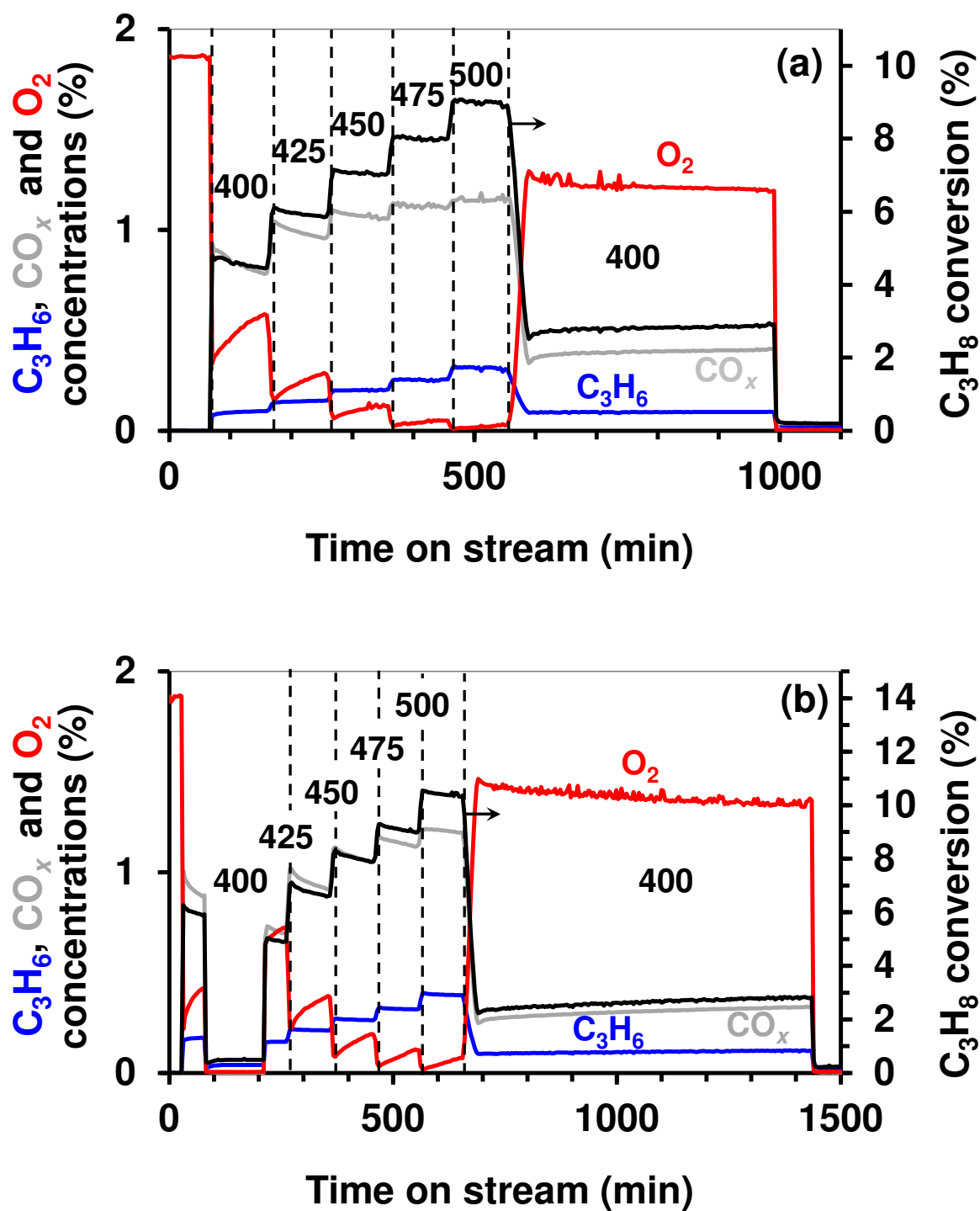
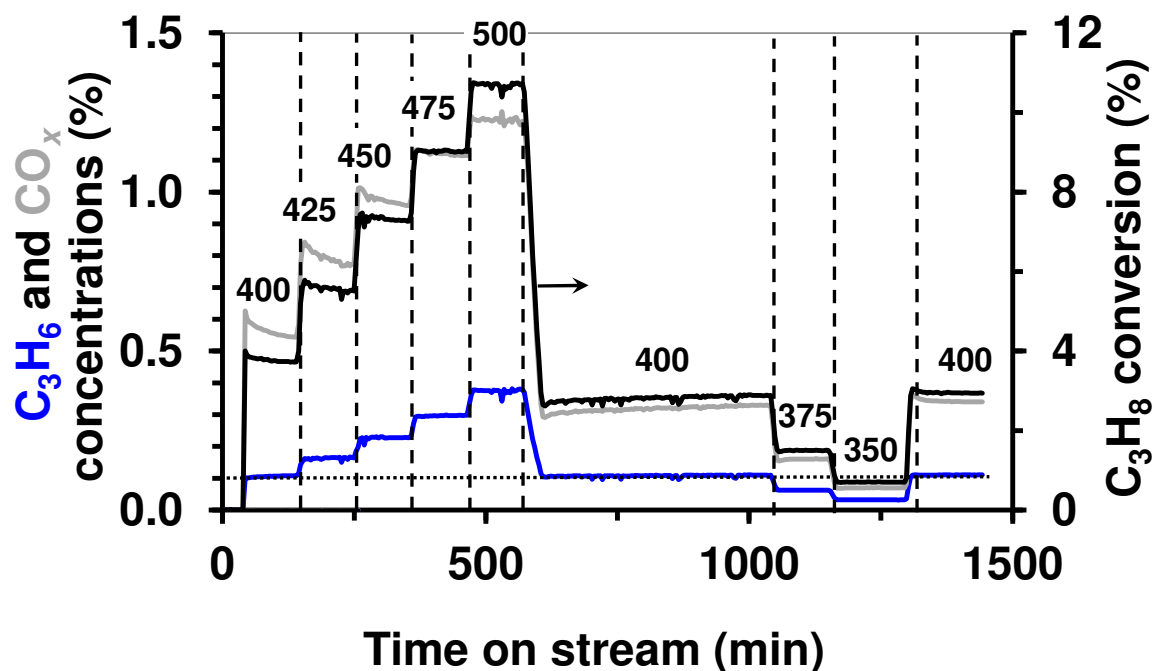


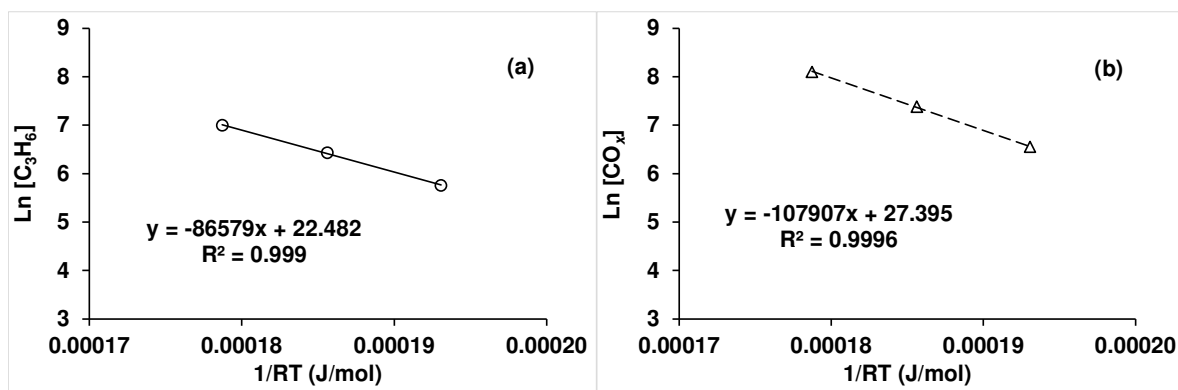
Figure S6. Influence of the removal of  $O_2$  from a 8.2%  $C_3H_8$  (8.2%) –  $O_2$  (1.9%) – He feed at 400 °C after ODHP reaction from 400 to 500 °C on (a)  $Co_{3.0}$ -Si $\beta$ -pH-2.5 and (b)  $Co_{3.1}$ -Si $\beta$ -pH-



9.0 on the propane conversion (black trace) and  $\text{CO}_x$  (gray trace),  $\text{C}_3\text{H}_6$  (blue trace) and  $\text{O}_2$  (red trace) concentrations.



**Figure S7.** Evolution of the conversion of  $\text{C}_3\text{H}_8$  (black trace) and the concentrations in  $\text{C}_3\text{H}_6$  (blue trace) and  $\text{CO}_x$  (gray trace) (8.2%  $\text{C}_3\text{H}_8$ , 1.9%  $\text{O}_2$  in He, 74.1  $\text{mL}_{\text{NTP}} \text{min}^{-1}$  total flow) as a function of time on stream and reacting temperature on 60 mg of  $\text{Co}_{1.9}\text{-Si}\beta\text{-pH-2.5}$  diluted in 600 mg of  $\alpha\text{-Al}_2\text{O}_3$ . The numbers reported on the figures are related to the reaction temperature.



**Figure S8.** Arrhenius plots for the formation of (a) C<sub>3</sub>H<sub>6</sub> and (b) CO<sub>x</sub> over 60 mg of Co<sub>1.9</sub>-Siβ-pH-2.5 diluted in 600 mg of α-Al<sub>2</sub>O<sub>3</sub> in the 350-400 °C temperature range (8.2% C<sub>3</sub>H<sub>8</sub>, 1.9% O<sub>2</sub> in He, 74.1 mL<sub>NTP</sub> min<sup>-1</sup> total flow).