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Article

# Thermochemical Activity of Single- and Dual-Phase Oxide Compounds Based on Ceria, Ferrites, and Perovskites for Two-Step Synthetic Fuel Production

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**Abstract:** This study focuses on the generation of solar thermochemical fuel (hydrogen, syngas) from CO<sub>2</sub> and H<sub>2</sub>O molecules via two-step thermochemical cycles involving intermediate oxygen-carrier redox materials. Different classes of redox-active compounds based on ferrite, fluorite, and perovskite oxide structures are investigated, including their synthesis and characterization associated with experimental performance assessment in two-step redox cycles. Their redox activity is investigated by focusing on their ability to perform the splitting of CO<sub>2</sub> during thermochemical cycles while quantifying fuel yields, production rates, and performance stability. The shaping of materials as reticulated foam structures is then evaluated to highlight the effect of morphology on reactivity. A series of single-phase materials including spinel ferrite, fluorite, and perovskite formulations are first investigated and compared to state-of-the-art materials. NiFe<sub>2</sub>O<sub>4</sub> foam exhibits a CO<sub>2</sub>-splitting activity similar to its powder analog after reduction at 1400 °C, surpassing the performance of ceria but with much slower oxidation kinetics. On the other hand, although identified as high-performing materials in other studies, Ce<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>2</sub>, Ca<sub>0.5</sub>Ce<sub>0.5</sub>MnO<sub>3</sub>, Ce<sub>0.2</sub>Sr<sub>1.8</sub>MnO<sub>4</sub>, and Sm<sub>0.6</sub>Ca<sub>0.4</sub>Mn<sub>0.8</sub>Al<sub>0.2</sub>O<sub>3</sub> are not found to be attractive candidates in this work (compared with La<sub>0.5</sub>Sr<sub>0.5</sub>Mn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub>). In the second part, characterizations and performance evaluation of dual-phase materials (ceria/ferrite and ceria/perovskite composites) are performed and compared to single-phase materials to assess a potential synergistic effect on fuel production. The ceria/ferrite composite does not provide any enhanced redox activity. In contrast, ceria/perovskite dual-phase compounds in the form of powders and foams are found to enhance the CO<sub>2</sub>-splitting performance compared to ceria.

**Keywords:** hydrogen; water-splitting; CO<sub>2</sub> conversion; thermochemical cycles; concentrated solar energy; perovskite; oxygen-carrier redox materials; solar fuels; reticulated foams; composites



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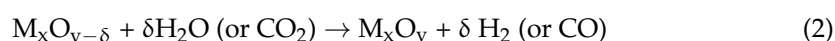


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## 1. Introduction

In the frame of the production of solar thermochemical fuels, two-step thermochemical cycles for splitting water or CO<sub>2</sub> molecules offer several potential advantages. The first is the high theoretical energy conversion efficiency that can be achieved through the direct use of concentrated solar energy without the intermediate solar heat or photons conversion to electricity [1]. Solar thermochemical fuel production is indeed advantageous compared to electrochemical or photo(electro)chemical routes because it uses the entire solar spectrum for splitting H<sub>2</sub>O and CO<sub>2</sub> molecules and, as such, provides a thermodynamically favorable path for efficient solar energy conversion and storage into fuels. Oxygen release and H<sub>2</sub>/CO production occur in different steps, avoiding their possible recombination and high-temperature gas separation. Moreover, the need for noble and costly metal catalysts is not required, and all the oxygen-carrier materials involved in the process are recycled.

H<sub>2</sub>O and CO<sub>2</sub> are the only inputs while concentrated solar energy is used as the high-temperature heat source for thermochemical reactions. Two-step thermochemical cycles involve a redox couple (M<sub>x</sub>O<sub>y</sub>/M<sub>x</sub>O<sub>y-δ</sub>), which allows the temperature to be reduced compared to the direct thermolysis of H<sub>2</sub>O or CO<sub>2</sub> molecules. In the first step (Equation (1)), an oxide is thermally reduced under highly concentrated solar energy, thereby creating oxygen vacancies and releasing oxygen from the oxide lattice (at ~1400 °C). During the second step (Equation (2)), the activated oxide reacts with water vapor or CO<sub>2</sub> to be re-oxidized and produce H<sub>2</sub> or CO.



Different types of relevant metal oxide redox cycles have been considered. They chiefly include volatile (such as ZnO/Zn, SnO<sub>2</sub>/SnO, MgO/Mg) or non-volatile oxides (such as Fe<sub>3</sub>O<sub>4</sub>/FeO and ferrites, CeO<sub>2</sub>/CeO<sub>2-δ</sub> and substituted ceria, and ABO<sub>3</sub>/ABO<sub>3-δ</sub> perovskites systems). The first category involves simple oxides with discrete transitions in the metal oxidation state (leading to high oxygen amounts being released and recovered during each cycle) and thus produces high fuel amounts per mass of active material at the expense of recombination issues with O<sub>2</sub> during the reduction step. The second category (that chiefly involves non-stoichiometric oxides) does not exhibit any phase change while O<sub>2</sub> is released from the solid, thus limiting the recombination and avoiding the need for the oxide transport between steps. However, they feature a lower mass-specific fuel productivity than the stoichiometric oxides due to a lower amount of oxygen being exchanged during solid-state reactions (quantified by the magnitude of oxygen non-stoichiometry δ). Currently, three main classes of redox-active non-stoichiometric compounds are considered to act as intermediate redox materials. Ferrites, ceria, and perovskite-based oxides have the ability to be partially reduced and oxidized reversibly without any phase changes (solid state) during redox reactions, maintaining their stable crystallographic structure. They feature a continuous redox activity over changes in temperature and oxygen partial pressure, thanks to the creation of oxygen vacancies that facilitate ionic diffusion. Nickel ferrites (NiFe<sub>2</sub>O<sub>4</sub>) were first identified as the most promising iron-based oxides for thermochemical cycles [2,3]. A large reduction yield was achieved at 1400 °C, and a good water-splitting capacity of the reduced species was demonstrated. The main drawback of this oxide is the low thermal stability after several cycles. The material tends to sinter and in turn, the reactivity decreases over cycles. Several solutions have been considered to improve thermal stability, such as supporting Ni-ferrites on ZrO<sub>2</sub> support [4,5] or using ferrite powder in a fluidized bed reactor to limit particle growth and promote heat and mass transfer [6]. Ceria and doped ceria have also been widely investigated for use in thermochemical redox cycles based on an initial study in 2006 [7–10]. Pure ceria demonstrated fast redox kinetics and good thermal stability during cycling, but the process is limited by a relatively low fuel productivity per gram of material due to the low reduction extent δ. High temperatures and/or low partial pressures of O<sub>2</sub> during the reduction step are additionally required to enhance δ. Alternatively, ceria doping with cations (such as Zr<sup>4+</sup>) can increase the reduction efficiency and fuel productivity at the expense of a decrease in oxidation kinetics [11,12]. Another way to improve material reactivity and kinetics is to use materials with designed morphologies and microstructures [13], such as reticulated porous structures [14–16], 3D-ordered macroporous structures [17,18], fibers [19,20], and porous microspheres [21]. A third family of oxides actively studied consists of the perovskite oxides for their capacity to be thermally reduced while maintaining their crystalline structure stable and to adapt to a large space of compositions. This great versatility permits to design and optimize materials with improved thermochemical performance. Manganite perovskites such as La<sub>0.5</sub>Sr<sub>0.5</sub>MnO<sub>3</sub> [22,23], BaCe<sub>0.25</sub>Mn<sub>0.75</sub>O<sub>3</sub> [24], or La<sub>0.8</sub>Sr<sub>0.2</sub>(Mn<sub>0.2</sub>Fe<sub>0.2</sub>Co<sub>0.4</sub>Al<sub>0.2</sub>)O<sub>3</sub> [25] have shown attractive fuel production yields per cycle although they are still limited by kinetics.

Although these materials have shown their ability to be used in two-step solar thermochemical cycles, an improvement in the overall solar-to-fuel efficiency is still expected. Improving the performance of materials is the main lever for success among the different possibilities such as the tuning of process conditions or the optimization of reactor design [26,27]. Indeed, even a small increase in material productivity will have a strong impact over thousands of cycles day after day.

Most recent studies have focused on materials with spinel, perovskite, and fluorite structures as redox catalysts (oxygen carriers) in this process [28]. To date, very little research has considered a comparison of the thermochemical performance of redox materials based on such structures. In addition, the activity of dual-phase compounds has been scarcely studied so far. For this purpose, experimental studies are required to evaluate the reactivity of various compounds applied to thermochemical cycles and to identify the most suitable compositions. In this work, different formulations of redox-active materials were synthesized, characterized, and tested during thermochemical cycles based on various structural configurations and cationic substitutions in the crystal lattice of the materials, with the aim of comparing their performance with state-of-the-art materials. Additionally, their redox activity has been studied by focusing on their ability to perform the splitting of CO<sub>2</sub> or H<sub>2</sub>O molecules during thermochemical cycles while quantifying the fuel production yields, rates, and performance stability.

The thermochemical activity and stability of materials with different structures and formulations were compared under the same cycling conditions. The evolution of the reactivity from one cycle to another was discussed. Different materials' morphologies were also considered to show possible effects on fuel production yields.

First, the performance of nickel ferrite and ceria shaped as reticulated porous structures (open-cell foams) was compared, as they represent the state-of-the-art materials for this application. Then, other formulations were considered based on recent studies reported in the literature. Previous works indeed identified different new formulations such as Ce<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>2</sub>, Ca<sub>0.5</sub>Ce<sub>0.5</sub>MnO<sub>3</sub>, Ce<sub>0.2</sub>Sr<sub>1.8</sub>MnO<sub>4</sub>, and Sm<sub>0.6</sub>Ca<sub>0.4</sub>Mn<sub>0.8</sub>Al<sub>0.2</sub>O<sub>3</sub> as attractive candidates for two-step redox cycles, but some of them reported unusual performance in comparison to benchmark materials, such as ceria, ferrites, or lanthanum manganite perovskites. It is therefore necessary to carry out additional experimental tests to check whether such materials are really promising. These materials were thus selected in this work to assess and compare their redox activity under given operating conditions commonly used in thermochemical cycles and which are typically required for the splitting of the CO<sub>2</sub> molecule. Finally, dual-phase materials consisting of ceria/ferrite and ceria/perovskite composites were considered to determine whether adding a phase more reducible than ceria can improve overall fuel production.

## 2. Results and Discussion

### 2.1. Ni-Ferrite (NiFe<sub>2</sub>O<sub>4</sub>) Foam

NiFe<sub>2</sub>O<sub>4</sub> has been widely studied for application in two-step solar thermochemical cycles [29–31], and this work investigates the reactivity of nickel ferrite foam in order to compare it to ferrite powder and ceria foam. The nickel ferrite foam was obtained by the replication method detailed in the Materials and Methods section. Figure 1 shows a picture of the foam after calcination at 1400 °C (before redox cycles).

The ferrite foam was characterized by X-ray diffraction. The diffraction pattern of the crushed foam (before redox cycles) is shown in Figure 2. The spinel structure associated with the doped ferrite is observed.



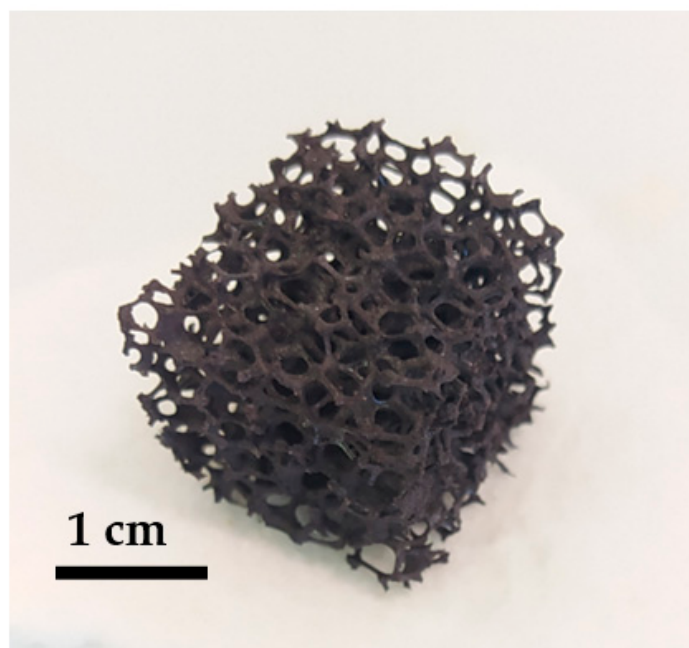


Figure 1. Picture of a pristine NiFe<sub>2</sub>O<sub>4</sub> foam calcined at 1400 °C.

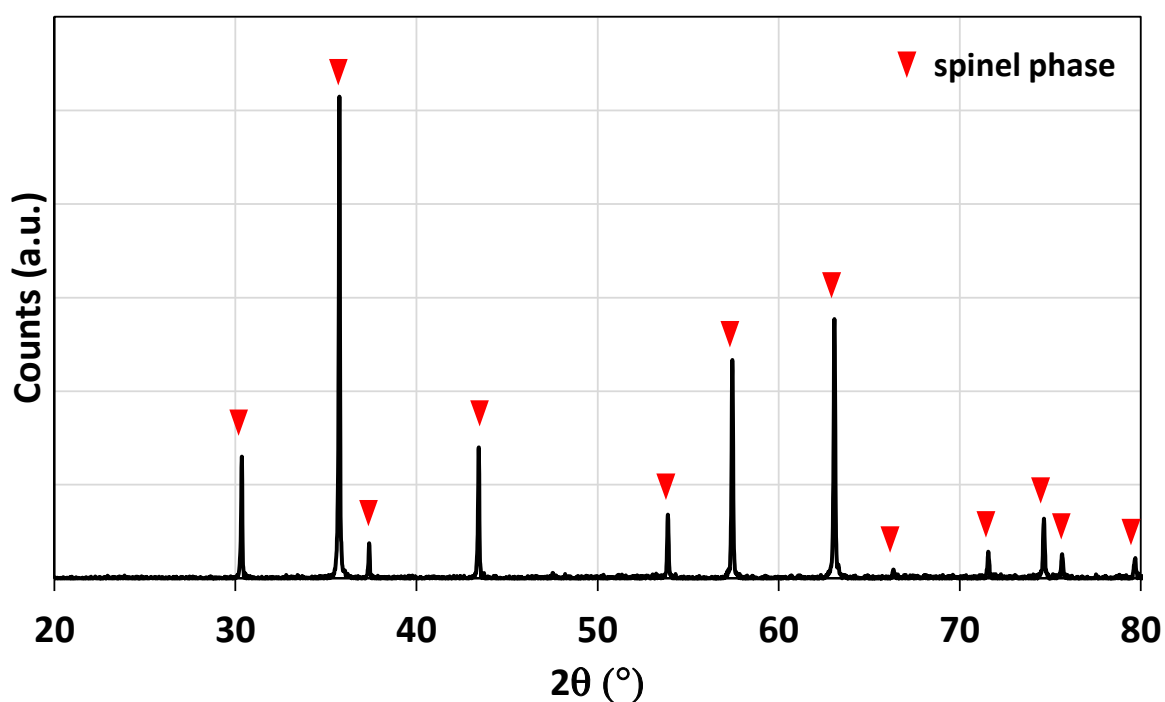
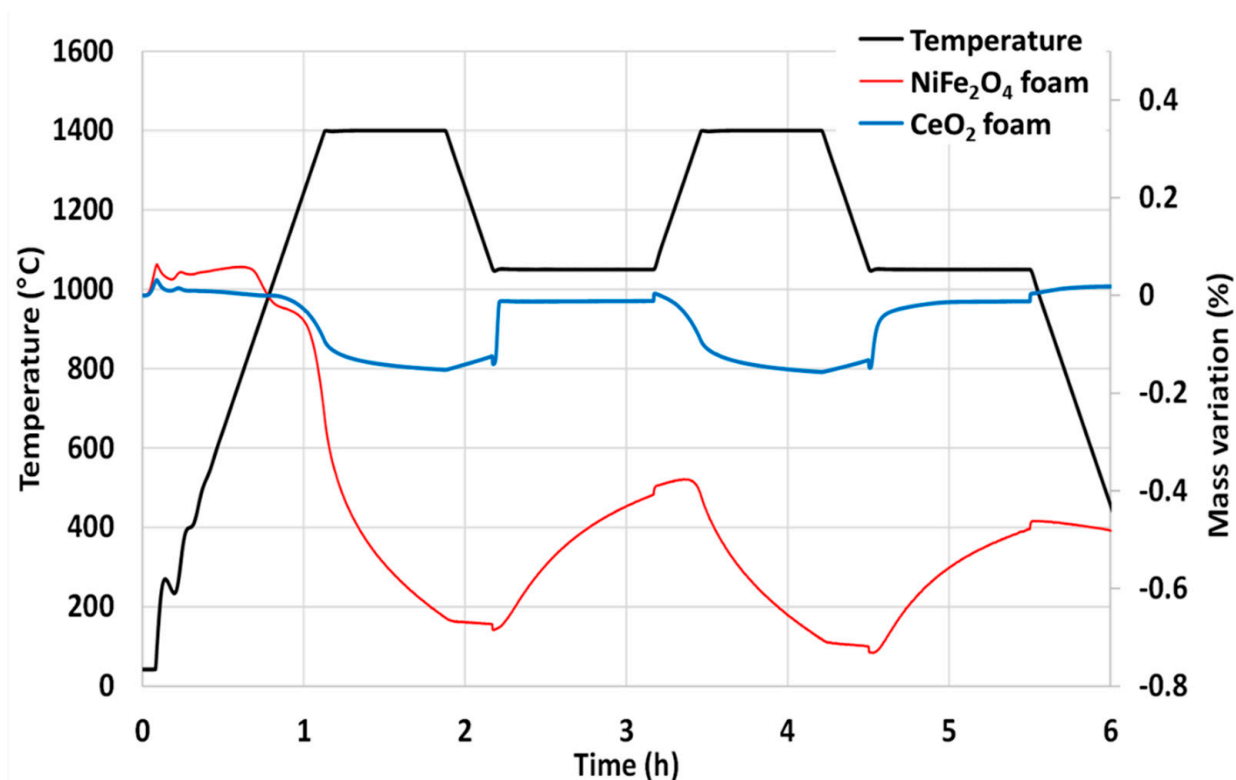


Figure 2. X-ray diffraction pattern of the NiFe<sub>2</sub>O<sub>4</sub> foam after calcination at 1400 °C.

The nickel ferrite foam was also characterized by thermogravimetric analysis (TGA) during two consecutive CO<sub>2</sub>-splitting cycles and further compared with ceria foam (Figure 3). The mass loss of the sample is due to the oxygen released from the material during the thermal reduction. The increase in mass is due to the oxygen uptake by the material during the supply of CO<sub>2</sub> (oxidant). The negative value of mass loss means that the mass of the sample is decreased from its initial mass due to the release of oxygen. It was found that 201 μmol of O<sub>2</sub> per gram of NiFe<sub>2</sub>O<sub>4</sub> was released during the first reduction step at 1400 °C and 172 μmol of CO per gram was produced during the first oxidation step with a peak production rate of 6.2 μmol<sub>CO</sub>/min.g (measured a few seconds after CO<sub>2</sub>

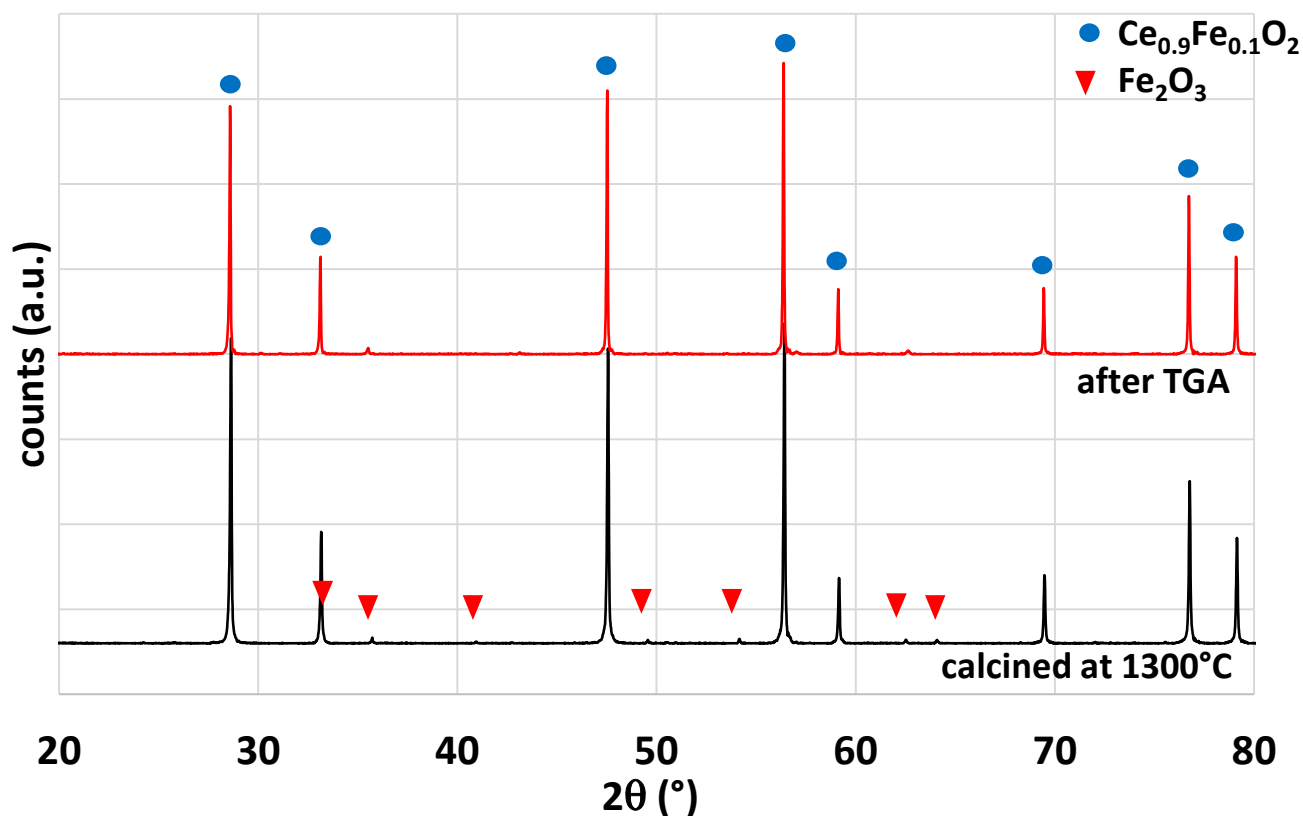
injection). The second cycle produced 106  $\mu\text{mol}$  of  $\text{O}_2$  per gram during the reduction step and 158  $\mu\text{mol}$  of  $\text{CO}$  per gram during the oxidation step with a peak production rate of 6.2  $\mu\text{mol}_{\text{CO}}/\text{min.g}$ . These values are consistent with previously reported data [32]. Compared to the  $\text{CeO}_2$  foam yielding about 100  $\mu\text{mol}_{\text{CO}}/\text{g}$  under the same conditions, the nickel ferrite foam produces much more  $\text{CO}$  but a slight decrease in reactivity is observed during cycling, contrary to ceria. This well-known phenomenon is explained by the limitations due to sintering and kinetic effects for ferrites, while their shaping as a porous foam does not alleviate this phenomenon.



**Figure 3.** Thermogravimetric analysis of two consecutive  $\text{CO}_2$ -splitting cycles with nickel ferrite and ceria foams.

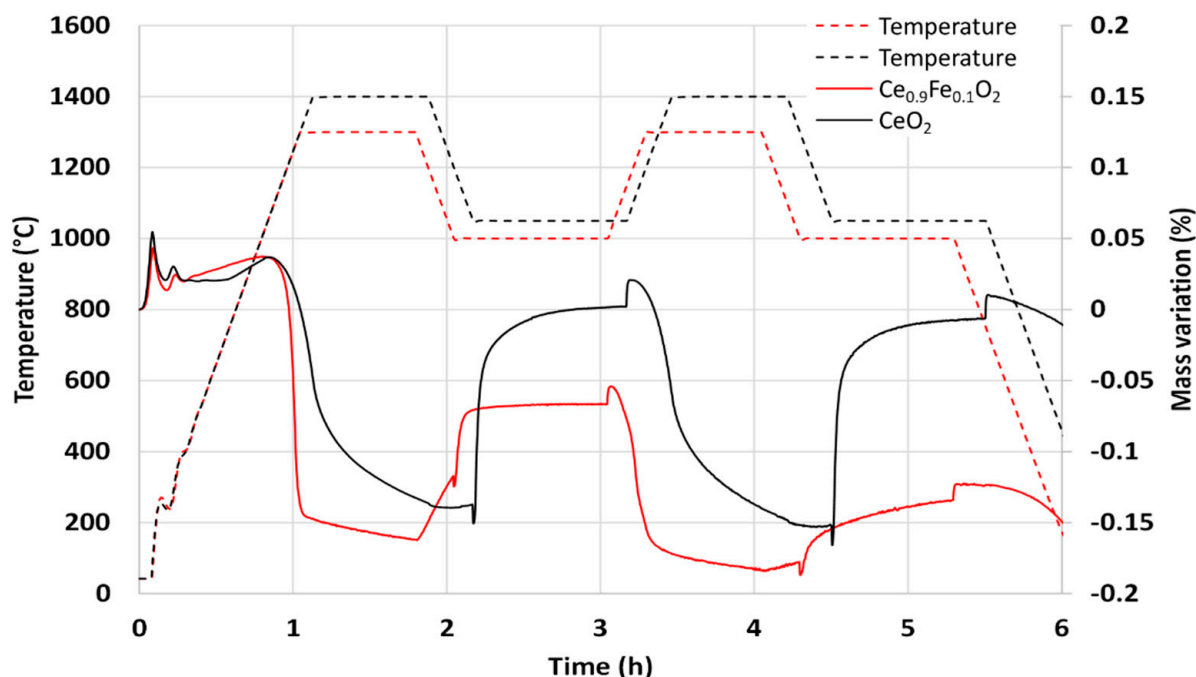
### 2.2. Iron-Doped Ceria ( $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$ )

Iron-doped ceria powder has been investigated during  $\text{CO}_2$ -splitting thermochemical cycles. This material was selected in this study because it was reported in the literature as a promising material for solar thermochemical fuel production [33,34].  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  was synthesized by the coprecipitation method detailed in the Materials and Methods section. The powder was calcined at 1300  $^\circ\text{C}$  for 1 h (this temperature was selected to alleviate the effect of sintering). On leaving the oven, the powder was partially sintered and thus ground in a mortar. The powder was characterized by XRD (Figure 4). Two phases were observed after calcination at 1300  $^\circ\text{C}$ : a major phase related to the crystallographic structure of ceria (fluorite) and a minor phase of  $\text{Fe}_2\text{O}_3$ . The cubic lattice parameter of the fluorite phase was calculated from the diffraction peaks and a value of  $a = 5.400$  (9)  $\text{\AA}$  was found. Compared to pure ceria ( $a = 5.4112$   $\text{\AA}$  [35]), a lower value was obtained for parameter  $a$ , which confirms the insertion of the iron atoms in the structure due to a smaller ionic radius. After cycles in TGA, the structure remained unchanged.



**Figure 4.** Comparison of the X-ray diffraction patterns for  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  after calcination at 1300 °C and after TGA analysis.

Two consecutive thermochemical cycles were performed in TGA. Figure 5 shows the mass variation during the  $\text{CO}_2$ -splitting cycles compared to pure ceria powder as a reference. The reduction temperature was 1300 °C (the same as the calcination temperature), and the oxidation step was carried out at 1000 °C. A steep mass loss was observed during the first reduction step associated with the reduction of  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$ , corresponding to an  $\text{O}_2$  release of 62  $\mu\text{mol/g}$ . The re-oxidation was not complete during the first cycle with a  $\text{CO}$  production of only 36  $\mu\text{mol/g}$  with a peak production rate of 15.1  $\mu\text{mol}_{\text{CO}}/\text{min.g}$ . During the second cycle, an  $\text{O}_2$  release of 40  $\mu\text{mol/g}$  and a  $\text{CO}$  production of 33  $\mu\text{mol/g}$  were measured with a peak production rate of 6.3  $\mu\text{mol}_{\text{CO}}/\text{min.g}$ . Compared to pure ceria, the reduction yield at 1300 °C was higher with faster kinetic rate, but the material undergoes strong sintering during thermal cycling, which leads to poor cyclability with a decrease in the reactivity. These results show different trends from those previously reported. Al-Taweel et al. [34] reported an  $\text{H}_2$  production 1.8 times greater than ceria after a reduction step at 1550 °C, while Orfila et al. [33] mentioned an  $\text{H}_2$  production of 8.5 STPcm<sup>3</sup>/g (380  $\mu\text{mol/g}$ ) with  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  powder after a reduction step at 1300 °C. The sintering issue was not reported as limiting the overall production performance of the cycle, and such high production yields are unexpected given the low thermal stability induced by iron doping. These results are therefore very different from each other and very far from the values found in this study.



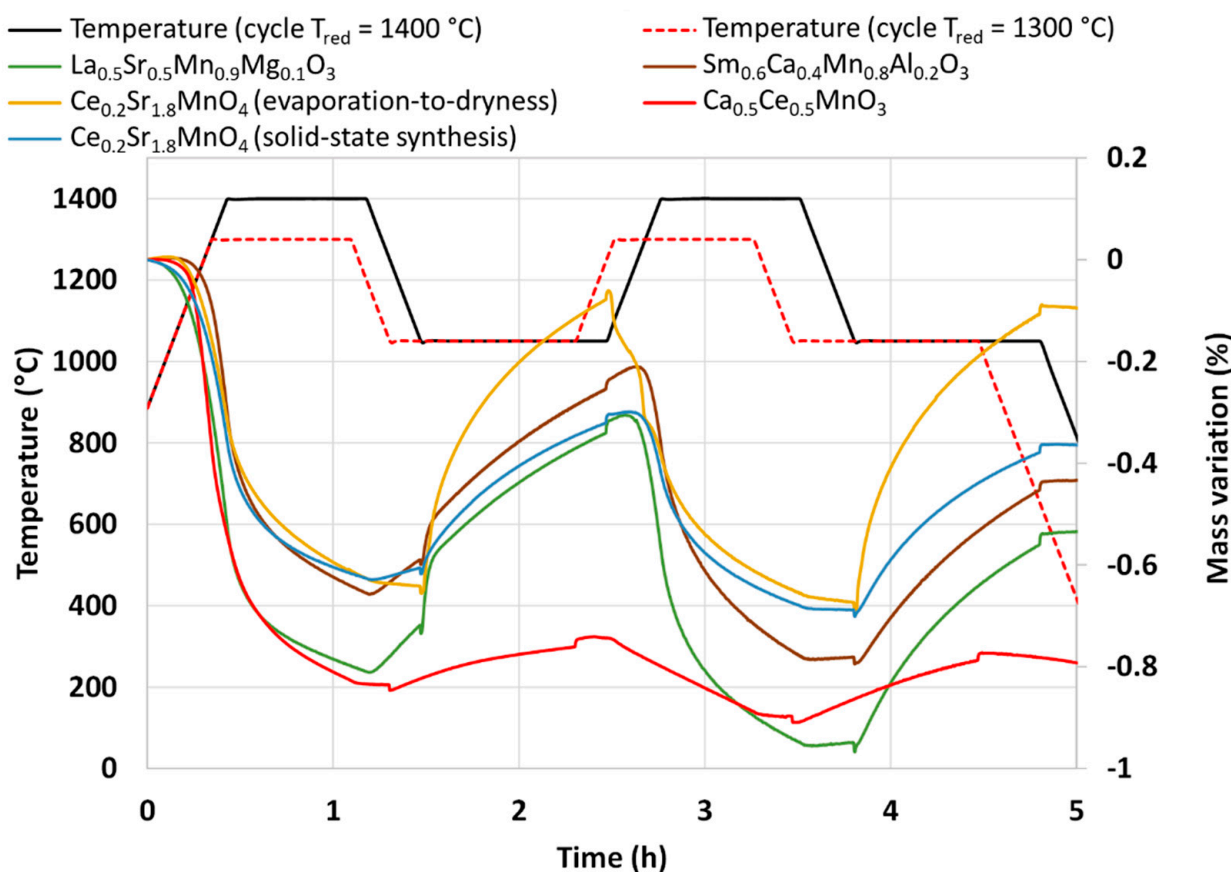
**Figure 5.** Thermogravimetric analysis of  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  during two consecutive  $\text{CO}_2$ -splitting thermochemical cycles. The TGA curve obtained for ceria is reported for comparison.

### 2.3. Perovskites

Several perovskite compositions were also synthesized and characterized during thermochemical cycles. Table 1 presents the synthesis method, the powder formulations, and their reactivity during  $\text{CO}_2$ -splitting cycles.  $\text{La}_{0.5}\text{Sr}_{0.5}\text{Mn}_{0.9}\text{Mg}_{0.1}\text{O}_3$  (LSMMg) is used as a reference material for the performance comparison as it was previously identified as a promising candidate material for such an application [36]. The thermogravimetric analyses are presented in Figure 6. Two consecutive  $\text{CO}_2$ -splitting cycles were carried out with a reduction temperature of  $1400\text{ }^\circ\text{C}$  (except for  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  which was reduced at  $1300\text{ }^\circ\text{C}$  due to a melting issue above this temperature) and a re-oxidation temperature of  $1050\text{ }^\circ\text{C}$ .

**Table 1.** Summary of  $\text{O}_2$  release and  $\text{CO}$  production yields, along with  $\text{CO}$  peak production rates during two consecutive thermochemical cycles with the considered series of perovskite powders (note that the reduction step for  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  was carried out at  $1300\text{ }^\circ\text{C}$ ).

Perovskite Composition	Synthesis Method	1st Cycle			2nd Cycle		
		$\text{O}_2$ Yield	$\text{CO}$ Yield	$\text{CO}$ Peak Production Rate	$\text{O}_2$ Yield	$\text{CO}$ Yield	$\text{CO}$ Peak Production Rate
		( $\mu\text{mol/g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/min.g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/min.g}$ )
$\text{La}_{0.5}\text{Sr}_{0.5}\text{Mn}_{0.9}\text{Mg}_{0.1}\text{O}_3$	Pechini	254	246	15.3	203	249	12.7
$\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$	Evaporation-to-dryness	201	219	-	114	224	-
$\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$	Mecano-synthesis	197	184	15.7	121	199	6.9
$\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$	Pechini	260 ( $1300\text{ }^\circ\text{C}$ )	53	1.6	48 ( $1300\text{ }^\circ\text{C}$ )	76	2.2
$\text{Sm}_{0.6}\text{Ca}_{0.4}\text{Mn}_{0.8}\text{Al}_{0.2}\text{O}_3$	Pechini	206	213	25.4	180	212	6.4



**Figure 6.** Thermogravimetric analyses of different perovskites during two consecutive CO<sub>2</sub>-splitting thermochemical cycles (note that the reduction step for Ca<sub>0.5</sub>Ce<sub>0.5</sub>MnO<sub>3</sub> was carried out at 1300 °C).

LSMMg released 254 and 203  $\mu\text{mol/g}$  of O<sub>2</sub> during the reduction steps and produced 246 and 249  $\mu\text{mol/g}$  of CO during the re-oxidation steps.

The Ruddlesden–Popper (RP) perovskite Ce<sub>0.2</sub>Sr<sub>1.8</sub>MnO<sub>4</sub> was identified by Barcellos et al. from thermodynamic calculations as a good candidate for the production of synthetic fuel via thermochemical cycles [37]. Bergeson-Keller et al. experimentally tested this material during water-splitting cycles [38]. In this work, Ce<sub>0.2</sub>Sr<sub>1.8</sub>MnO<sub>4</sub> was synthesized by a soft synthesis method (evaporation-to-dryness method) and a solid-state mechanical method both described in the Materials and Methods section. The material synthesized by the Pechini method produced 201 and 114  $\mu\text{mol/g}$  of O<sub>2</sub> and 219 and 224  $\mu\text{mol/g}$  of CO from mass losses measured by TGA. XRD analysis (Figure 7) reveals the presence of a SrO phase that could react with CO<sub>2</sub> to produce SrCO<sub>3</sub> [39]. This side reaction participates in the mass increase during the re-oxidation step, and this is confirmed during the second reduction step by a first mass decrease just after the dwell at 1050 °C when the CO<sub>2</sub> injection was stopped. This is attributed to the strontium carbonate dissociation releasing CO<sub>2</sub>. A second synthesis method was used to obtain the pure RP perovskite phase without a SrO secondary phase (solid-state synthesis), and the material was tested in the same way with the thermobalance. The XRD analysis confirms the absence of a SrO phase (Figure 8). Accordingly, the mass variations were lower during TGA. The O<sub>2</sub> release was 197 and 121  $\mu\text{mol/g}$  and the CO production was 184 and 199  $\mu\text{mol/g}$ . The previous study found in the literature [38] reported an O<sub>2</sub> release at 1400 °C of 168 and 142  $\mu\text{mol/g}$  and a hydrogen production of 247  $\mu\text{mol/g}$  at 1000 °C with a 40%<sub>vol</sub> steam concentration. The material obtained in this study thus released more O<sub>2</sub> during the first reduction, but the CO yield was hindered by kinetic limitations. Note that a longer duration for the re-oxidation step would result in higher fuel production, since the oxidation step with CO<sub>2</sub> did not reach

completion after 45 min at 1050 °C. This investigation also reveals the importance of the synthesis method on materials' performance.

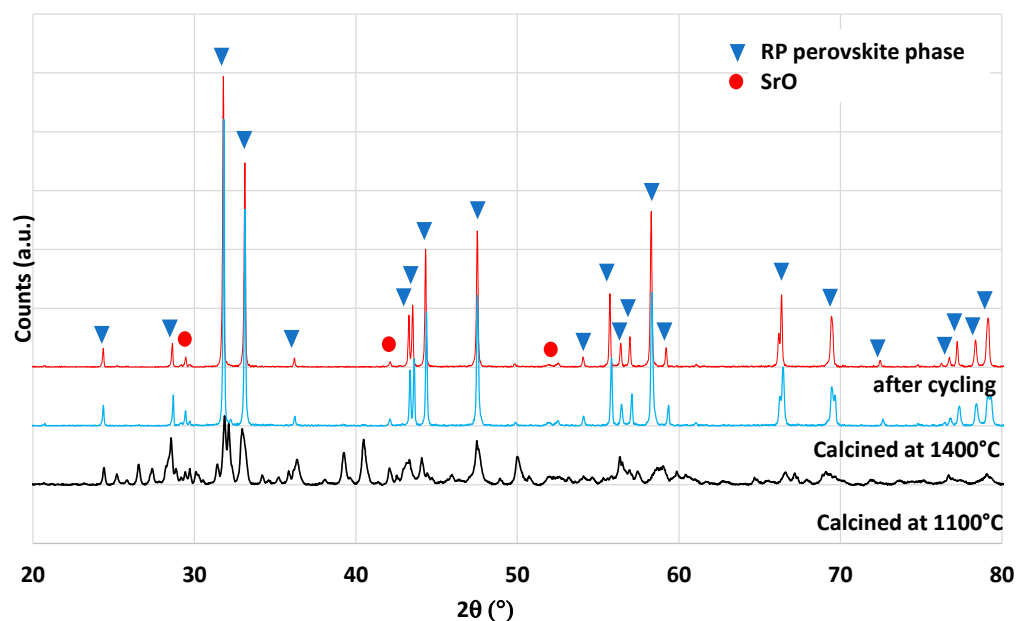


Figure 7. XRD pattern of  $\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$  obtained from the evaporation-to-dryness method.

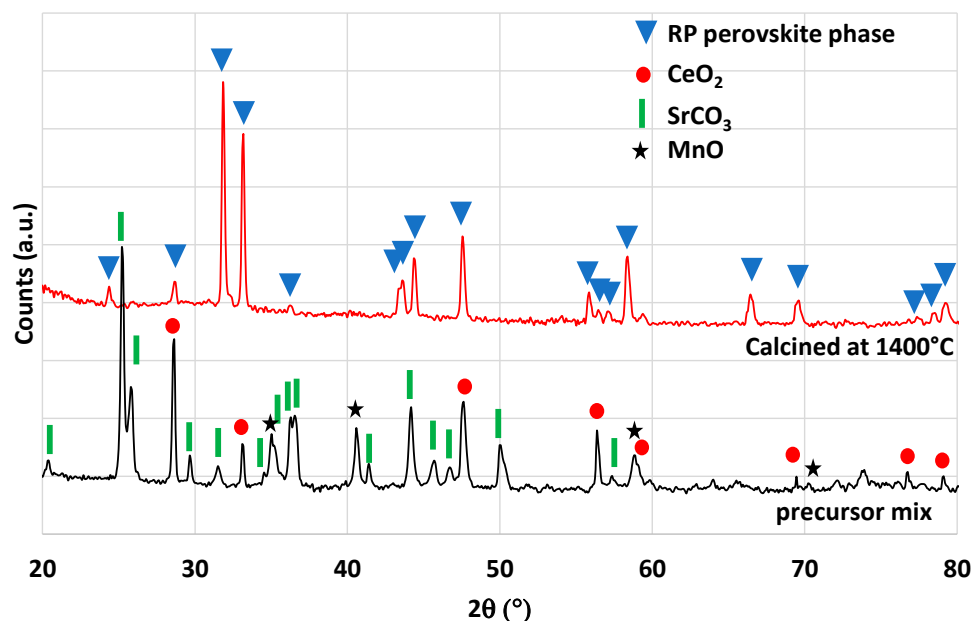
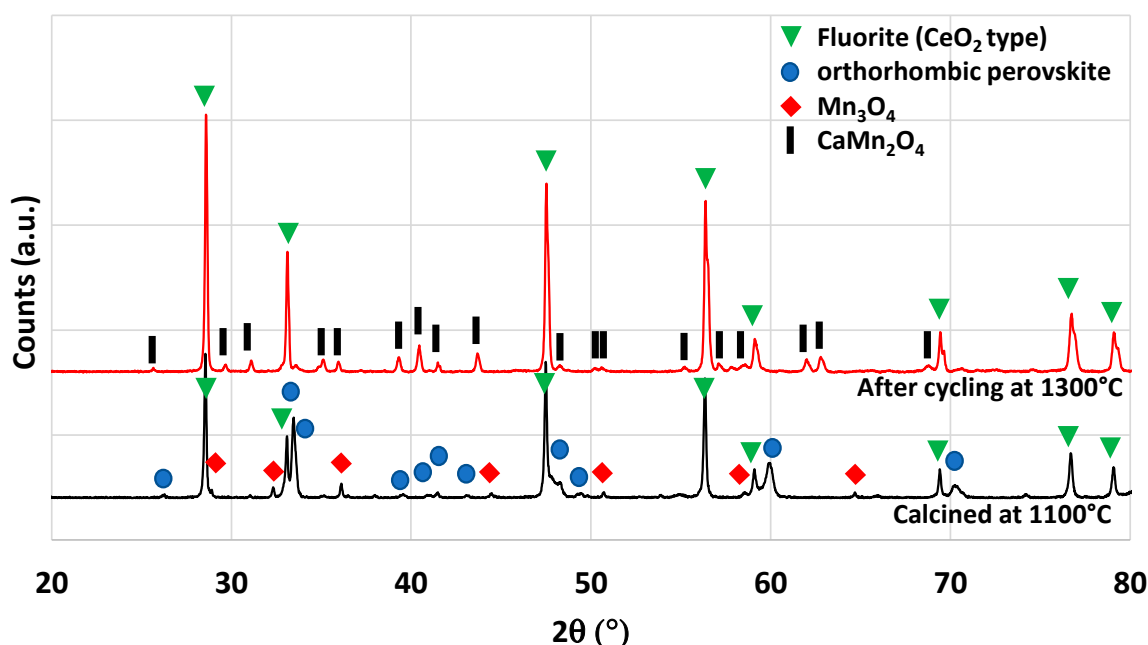


Figure 8. XRD patterns of  $\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$  obtained at 1400 °C by solid-state synthesis. The starting mixture of precursors is reported for comparison.

Another material formulation that was synthesized is  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$ . This perovskite composition was previously identified by theoretical calculations as a potential candidate with improved reactivity for thermochemical cycles [40]. The authors of this work mentioned the ability of simultaneous reduction of  $\text{Ce}^{4+}$  and  $\text{Mn}^{3+}$ . However, it was never tested experimentally to confirm the theoretical predictions. The material was thus synthesized by the modified Pechini method (see the Materials and Methods section). The XRD pattern of the synthesized powder is shown in Figure 9. After calcination at 1100 °C, two main phases are observed, the fluorite structure associated with  $\text{CeO}_2$  and the perovskite



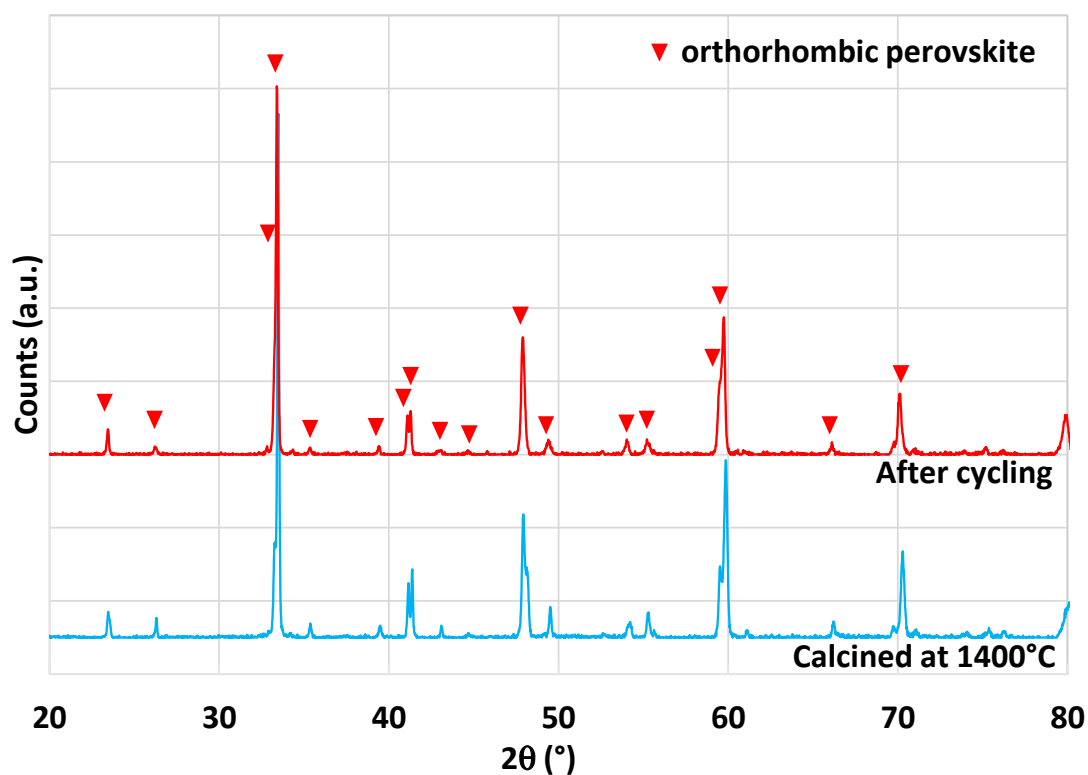
phase associated with  $(\text{Ca,Ce})\text{MnO}_3$ . A minor phase of  $\text{Mn}_3\text{O}_4$  is also observed on the XRD pattern. This observation is consistent with a previous study on  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  synthesis, where  $\text{CeO}_2$  was found to be a major phase obtained [41]. Hence, this material is a multi-phase material and the pure  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  phase was not obtained. The variation of mass during redox cycles from TG analysis is presented in Figure 6. Due to melting phenomena occurring at  $1400\text{ }^\circ\text{C}$ , the two consecutive cycles were performed with a reduction temperature of  $1300\text{ }^\circ\text{C}$ . During redox cycling,  $260\text{ }\mu\text{mol/g}$  and  $48\text{ }\mu\text{mol/g}$  of  $\text{O}_2$  were released, while  $53\text{ }\mu\text{mol/g}$  and  $76\text{ }\mu\text{mol/g}$  of  $\text{CO}$  were produced, respectively. At the end of the thermochemical cycles, the recovered powder was strongly sintered. Such performance does not overpass the gas production yields of ceria during  $\text{CO}_2$ -splitting cycles after a reduction step at  $1300\text{ }^\circ\text{C}$ . XRD analysis after thermochemical cycling is shown in Figure 9, which reveals the formation of a  $\text{CaMn}_2\text{O}_4$  phase instead of the perovskite phase that has totally disappeared after thermochemical cycles.



**Figure 9.** XRD patterns of  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  calcined at  $1100\text{ }^\circ\text{C}$  and after thermochemical cycling at  $1300\text{ }^\circ\text{C}$ .

Another perovskite formulation considered in this study is  $\text{Sm}_{0.6}\text{Ca}_{0.4}\text{Mn}_{0.8}\text{Al}_{0.2}\text{O}_3$ . This material was identified as highly promising during an experimental study [42], in which authors reported a very high and unusual  $\text{CO}$  production yield of  $580\text{ }\mu\text{mol/g}$  over 14 consecutive cycles with a reduction step at  $1350\text{ }^\circ\text{C}$ . Thus, a complementary analysis of this material was performed in this work to compare results and discuss the performance. The material was synthesized by the modified Pechini method (see the Section 3). XRD analysis is shown in Figure 10. The orthorhombic perovskite phase associated with the  $\text{Sm}_{0.6}\text{Ca}_{0.4}\text{Mn}_{0.8}\text{Al}_{0.2}\text{O}_3$  phase was identified after calcination at  $1400\text{ }^\circ\text{C}$ . After cycling, the perovskite phase remained stable. During the two consecutive  $\text{CO}_2$ -splitting cycles realized in the thermobalance, the  $\text{O}_2$  release was  $206$  and  $180\text{ }\mu\text{mol/g}$  and the  $\text{CO}$  production yield reached  $213$  and  $212\text{ }\mu\text{mol/g}$ . These values are thus far from those reported previously [42] and even lower than those of LSMMg used as reference formulation (Table 1). However, due to the slow kinetic rate, the oxidation step with  $\text{CO}_2$  did not reach completion after 45 min at  $1050\text{ }^\circ\text{C}$ , regardless of the considered perovskite material. Hence, a longer duration for the re-oxidation step would result in higher fuel production (the maximum  $\text{CO}$  production yield would however not exceed twice the  $\text{O}_2$  yield). Finally, this study confirms that the previously published unusual or unexpected data (e.g., related to  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  [33,34],  $\text{Ca}_{0.5}\text{Ce}_{0.5}\text{MnO}_3$  [40], or  $\text{Sm}_{0.6}\text{Ca}_{0.4}\text{Mn}_{0.8}\text{Al}_{0.2}\text{O}_3$  [42]) are

difficult to reproduce experimentally and therefore strongly require experimental validation and confirmation.



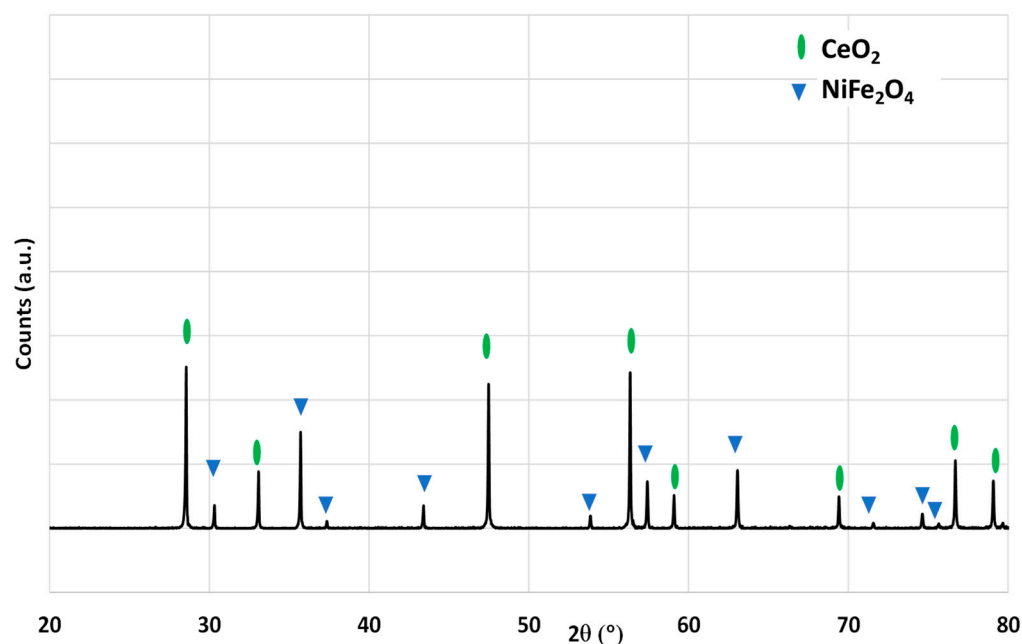
**Figure 10.** XRD patterns of  $\text{Sm}_{0.6}\text{Ca}_{0.4}\text{Mn}_{0.8}\text{Al}_{0.2}\text{O}_3$  prepared at 1400 °C and after thermochemical cycling at 1400 °C.

#### 2.4. Dual-Phase Composite Materials

Dual-phase materials were investigated to provide insights into the potential synergistic effect that may arise by mixing two redox-active materials with noticeable performance in thermochemical  $\text{CO}_2$ -splitting cycles. The activity of dual-phase compounds has been scarcely studied so far. Only a few articles reported the benefits of dual-phase materials containing ceria on the oxygen storage/exchange capacity [43–46]. In this study, the mixture of ceria with nickel ferrite was studied, as well as the mixture of ceria with a perovskite (LSMMg). These composite materials are expected to offer enhanced reactivity in two-step thermochemical cycles, compared to individual compounds.

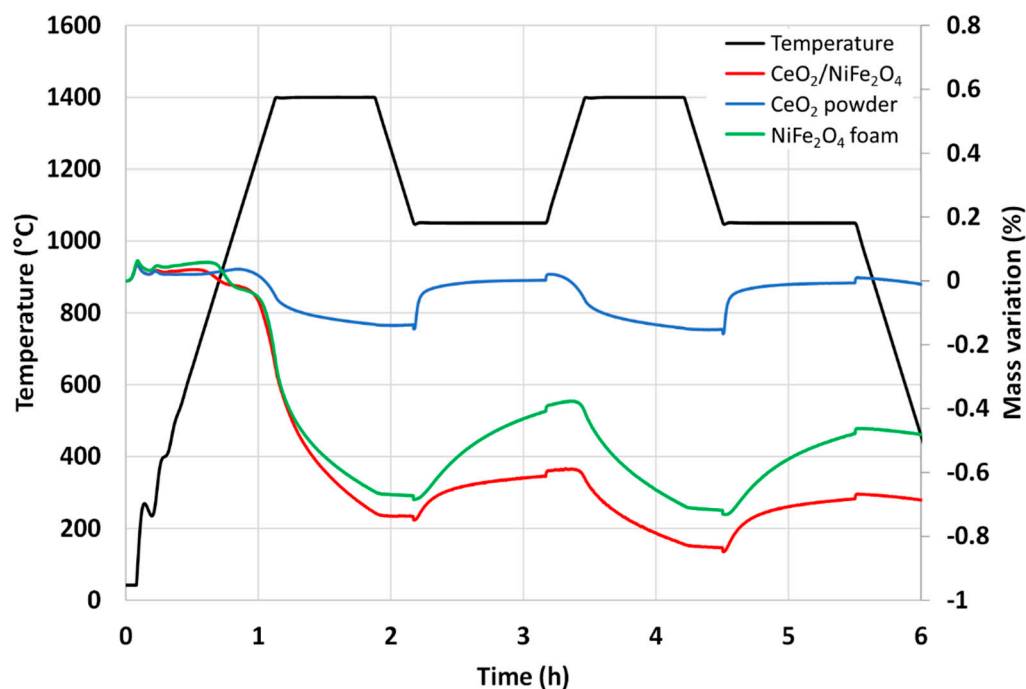
##### 2.4.1. Dual-Phase $\text{CeO}_2/\text{NiFe}_2\text{O}_4$

A mixture of ceria and nickel ferrite (50/50 wt%) was obtained by grinding a commercial micrometric powder of cerium oxide (Sigma-Aldrich, St. Louis, MO, USA, 99.9%, <5  $\mu\text{m}$  powder) with a previously synthesized nickel ferrite powder (obtained by the coprecipitation of hydroxides, see the Materials and Methods section) followed by calcination of the blend at 1400 °C under static air for 1 h. This thermal treatment permits to characterize the stability of dual-phase materials and to observe whether they react together at the temperature of the reduction step. The XRD pattern of the mixture after calcination is shown in Figure 11. The two distinct phases of ceria (fluorite) and nickel ferrite (spinel) are observed and no by-product has thus been formed, which confirms the thermal stability in air at 1400 °C.



**Figure 11.** XRD pattern of a mixture of  $\text{CeO}_2$  and  $\text{NiFe}_2\text{O}_4$  (50/50 wt%) after thermal treatment in air at  $1400\text{ }^\circ\text{C}$ .

The composite material was analyzed by TGA (Figure 12). During the two consecutive cycles, it released  $226\text{ }\mu\text{mol/g}$  and  $76\text{ }\mu\text{mol/g}$  of  $\text{O}_2$  during the reduction steps and produced  $85\text{ }\mu\text{mol/g}$  and  $104\text{ }\mu\text{mol/g}$  of  $\text{CO}$  during the oxidation steps (Table 2). Compared to the pure components, the dual-phase material did not show improved performance. The release of  $\text{O}_2$  was better than for pure ceria but lower than for pure Ni-ferrite, and the  $\text{CO}$  production yield was similar to that of ceria but lower than that of Ni-ferrite.

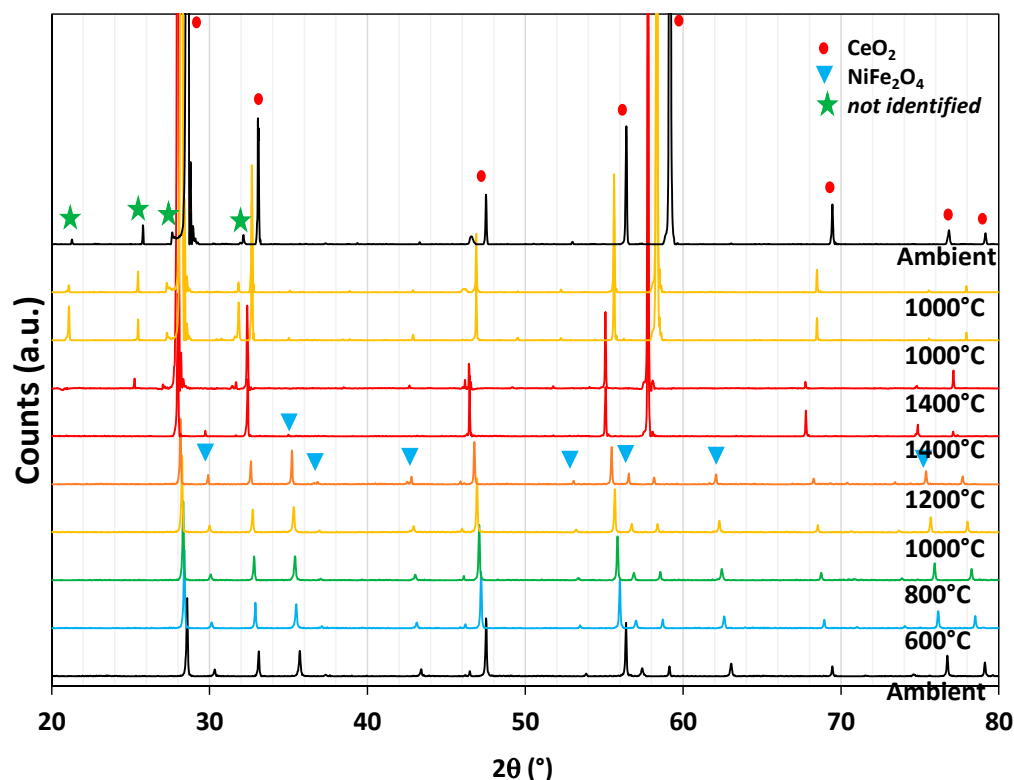


**Figure 12.** Thermogravimetric analysis of  $\text{CeO}_2/\text{NiFe}_2\text{O}_4$  mix (50/50 wt%) during two-consecutive  $\text{CO}_2$ -splitting thermochemical cycles, in comparison with single phases of  $\text{NiFe}_2\text{O}_4$  and  $\text{CeO}_2$ .

**Table 2.** O<sub>2</sub> and CO production yields, along with CO peak production rates during two-consecutive thermochemical cycles with the dual-phase material CeO<sub>2</sub>/NiFe<sub>2</sub>O<sub>4</sub> (50/50 wt%) compared to pristine ceria and Ni-ferrite.

Powder Composition	1st Cycle			2nd Cycle		
	O <sub>2</sub> Yield	CO Yield	CO Peak Production Rate	O <sub>2</sub> Yield	CO Yield	CO Peak Production Rate
	(μmol/g)	(μmol/g)	(μmol/min.g)	(μmol/g)	(μmol/g)	(μmol/min.g)
CeO <sub>2</sub> /NiFe <sub>2</sub> O <sub>4</sub> (50/50 wt%)	226	85	7.5	76	104	9.1
CeO <sub>2</sub>	55	95	39.5	54	98	43.7
NiFe <sub>2</sub> O <sub>4</sub>	201	172	6.2	106	158	6.2

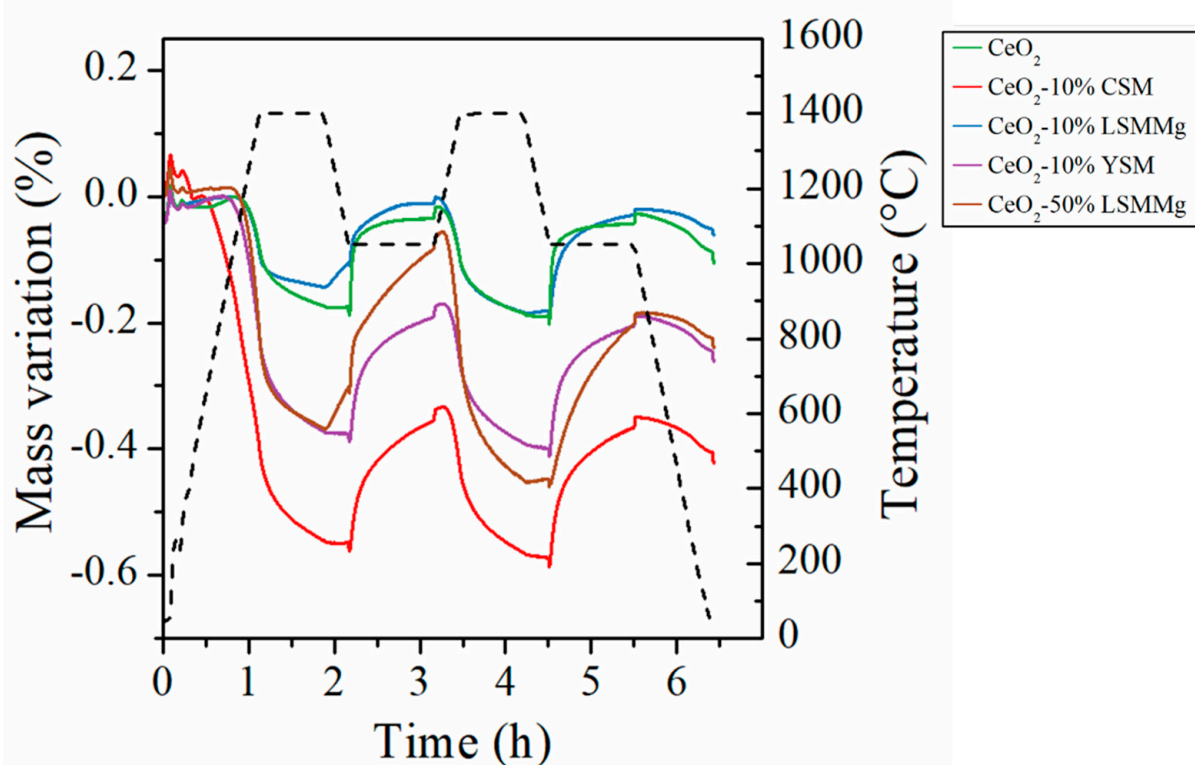
XRD analysis of the mixture at different temperatures under a controlled atmosphere is shown in Figure 13. At the bottom, the pattern was measured at room temperature. Going up from bottom to top, the powder was heated under a flow of nitrogen up to 1400 °C, with patterns acquisition at 600 °C, 800 °C, 1000 °C, 1200 °C, and 1400 °C. A plateau at 1400 °C for 1.5 h preceded a drop in temperature to 1000 °C at which pure CO<sub>2</sub> was swept inside the analysis chamber to react with the material. As a result, both ceria and Ni-ferrite phases are observed from room temperature to 1200 °C. The spinel phase corresponding to NiFe<sub>2</sub>O<sub>4</sub> disappears at 1400 °C, and a new phase not identified appears at the end of the reduction step. This means the ferrite phase mixed with ceria is not stable under inert gas at 1400 °C, whereas the same mixture was stable during thermal treatment in air (Figure 11). During cooling as well as during the re-oxidation step, the spinel phase does not reappear. The two materials composing the mixture thus react together during heating at 1400 °C under an inert atmosphere, and the reactivity is not improved. Therefore, the investigations on this composite material were stopped and another composite material was studied.



**Figure 13.** In situ XRD analysis of CeO<sub>2</sub>/NiFe<sub>2</sub>O<sub>4</sub> (50/50 wt%) during a thermochemical cycle under a controlled atmosphere (N<sub>2</sub> from ambient to 1400 °C and CO<sub>2</sub> at 1000 °C).

#### 2.4.2. Dual-Phase CeO<sub>2</sub>/Perovskite

Some studies have already mentioned the use of dual-phase ceria/perovskite materials in thermochemical cycles with expected benefits on fuel production performance [43,47]. The objective of this study on dual-phase materials is to observe whether a synergistic phenomenon can improve performance compared to single compounds, in particular by maintaining a good reduction yield (property of perovskites) combined with increased re-oxidation kinetics (property of ceria). In order to study the redox behavior of ceria/perovskite dual-phase compositions, TGA was performed with ceria–perovskite powder mixtures (Figure 14). Five mixture compositions were studied: ceria powder mixed with 10 wt% of La<sub>0.5</sub>Sr<sub>0.5</sub>Mn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> (LSMMg), 50 wt% of LSMMg, 10 wt% of Ca<sub>0.5</sub>Sr<sub>0.5</sub>MnO<sub>3</sub> (CSM), 50 wt% of CSM, and 10 wt% of Y<sub>0.5</sub>Sr<sub>0.5</sub>MnO<sub>3</sub> (YSM as an alternative material to CSM), abbreviated as CeO<sub>2</sub>-10% LSMMg, CeO<sub>2</sub>-50% LSMMg, CeO<sub>2</sub>-10% CSM, CeO<sub>2</sub>-50% CSM, and CeO<sub>2</sub>-10% YSM, respectively. The O<sub>2</sub> production yield was increased for CeO<sub>2</sub>-10% YSM, CeO<sub>2</sub>-50% LSMMg, CeO<sub>2</sub>-10% CSM, and CeO<sub>2</sub>-50% CSM in comparison with pristine ceria (Table 3). Accordingly, CO production yield was also increased, highlighting the beneficial impact on the redox activity of dual-phase compositions. Note that the CeO<sub>2</sub>-50%CSM material releases a large amount of O<sub>2</sub> (652 μmol/g in the first cycle) due to the large amount of CSM phase which is easily reduced (it is thus not represented in Figure 14). However, such a high reduction yield does not improve the CO production capacity of the material to a great extent because the re-oxidation step is not complete (151 and 168 μmol/g in the first and second cycles, respectively). The same behavior is observed for the CeO<sub>2</sub>-10% CSM material.

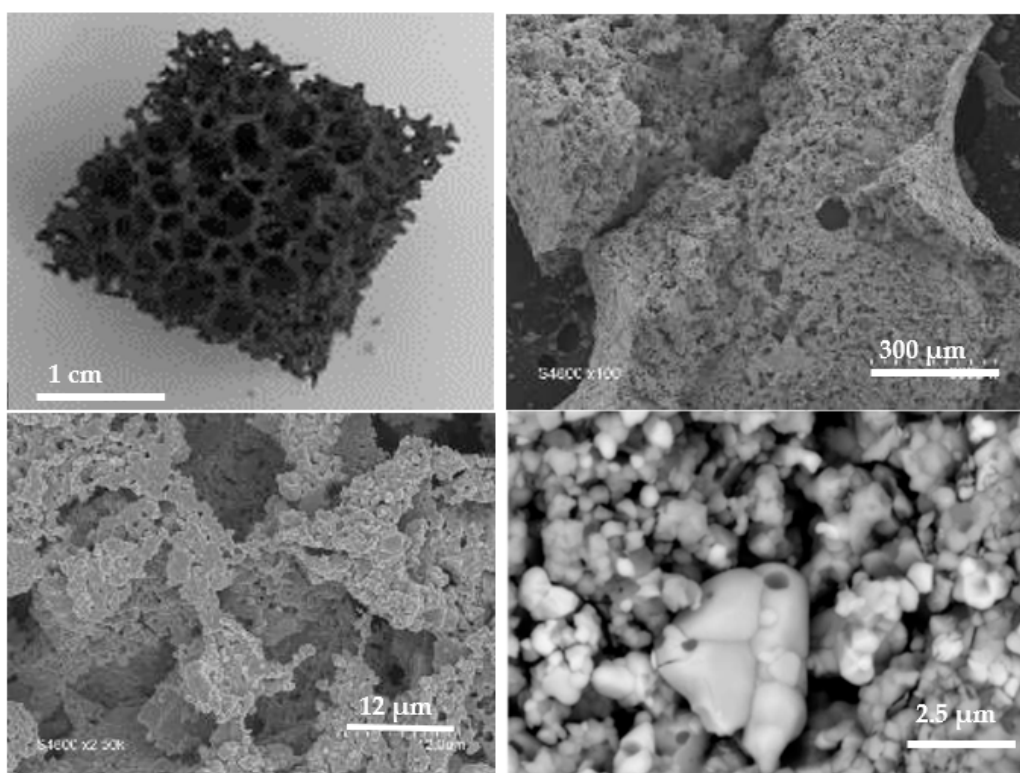


**Figure 14.** Thermogravimetric profiles (solid lines) along with the temperature profile (black dotted line) during thermochemical cycling of powder mixtures: CeO<sub>2</sub>-10% CSM, CeO<sub>2</sub>-10% LSMMg, CeO<sub>2</sub>-10% YSM, and CeO<sub>2</sub>-50% LSMMg, in comparison with single CeO<sub>2</sub>.

**Table 3.** O<sub>2</sub> and CO production yields, along with CO peak production rates during two thermochemical cycles in TGA for a series of dual-phase powder compositions, in comparison with single CeO<sub>2</sub> and LSMMg powders.

Powder Composition	1st Cycle			2nd Cycle		
	O <sub>2</sub> Yield	CO Yield	CO Peak Production Rate	O <sub>2</sub> Yield	CO Yield	CO Peak Production Rate
	( $\mu\text{mol/g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/min.g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/g}$ )	( $\mu\text{mol/min.g}$ )
CeO <sub>2</sub>	55	95	39.5	54	98	43.7
LSMMg	254	246	15.3	203	249	12.7
CeO <sub>2</sub> -10% YSM	117	123	15.0	71	129	16.7
CeO <sub>2</sub> -10% LSMMg	45	83	14.1	57	98	12.3
CeO <sub>2</sub> -50% LSMMg	120	178	25.7	124	163	6.6
CeO <sub>2</sub> -10% CSM	172	129	13.1	74	137	15.5
CeO <sub>2</sub> -50% CSM	652	151	5.1	93	168	5.8

In order to optimize the reactivity of the most interesting dual-phase material, a reticulated foam of CeO<sub>2</sub>/LSMMg (50/50 wt%) was prepared by the impregnation method after calcination of the powder mixture at 1400 °C in air (see the Materials and Methods section). Figure 15 presents the macroscopic and microscopic images of the foam. The foam exhibits a dark gray color due to LSMMg. The strands are highly porous and have a thickness of about 1 mm. Pores around grains are observed and the two materials are well mixed with a micrometric grain size. The foam thus exhibits a dual-scale porosity (millimetric cells and submicrometric pores within struts).



**Figure 15.** Macroscopic and microscopic images of the CeO<sub>2</sub>/LSMMg (50/50 wt%) dual-phase foam after calcination at 1400 °C.

Actually, the specific surface area of the synthesized materials is always low after their thermal treatment at 1400 °C ( $<1 \text{ m}^2/\text{g}$ ), which means that this property is not the main one impacting the CO yield. The porous structure of the foam can be advantageous because the most important is to keep good access of the reactive gas (CO<sub>2</sub>) to the bulk material



through interconnected porosity. Thus, sintering phenomena must be avoided as much as possible to maintain sufficiently high oxidation rates and thereby high CO yields.

The EDS analysis (Figure 16) shows that the two distinct phases remain, but it also reveals that the magnesium has migrated from the perovskite phase to form segregated spots on the ceria grains.

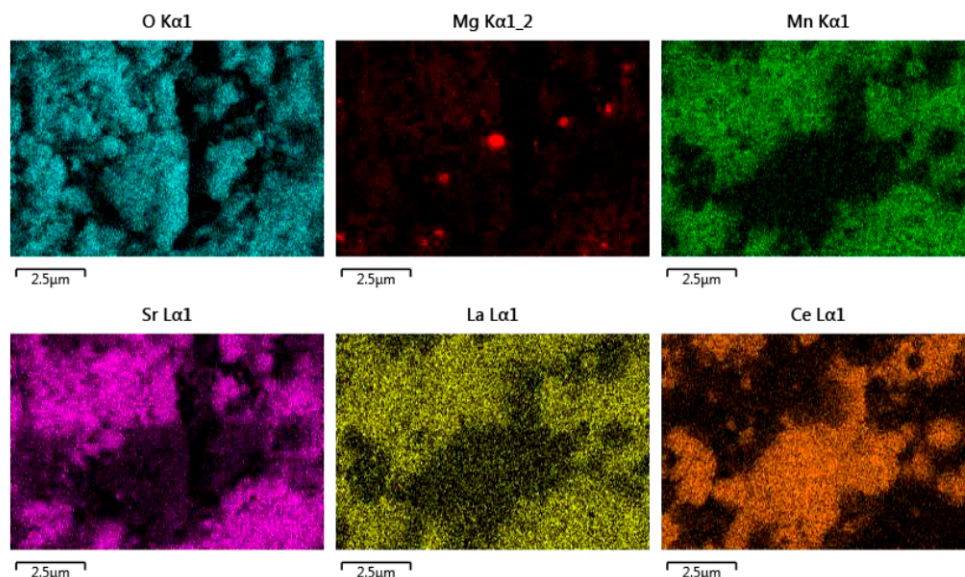


Figure 16. EDS analysis of  $\text{CeO}_2/\text{LSMMg}$  (50/50 wt%) dual-phase foam after calcination at 1400 °C.

The  $\text{CeO}_2/\text{LSMMg}$  foam was analyzed by TGA during two consecutive cycles (Figure 17). The releases of oxygen during the reduction steps are 173  $\mu\text{mol/g}$  and 96  $\mu\text{mol/g}$  and the CO production yields reach 176  $\mu\text{mol/g}$  and 180  $\mu\text{mol/g}$ . The production of CO with the foam is thus consistent with the production obtained with the mixture of powders without shaping under the same cycling conditions.

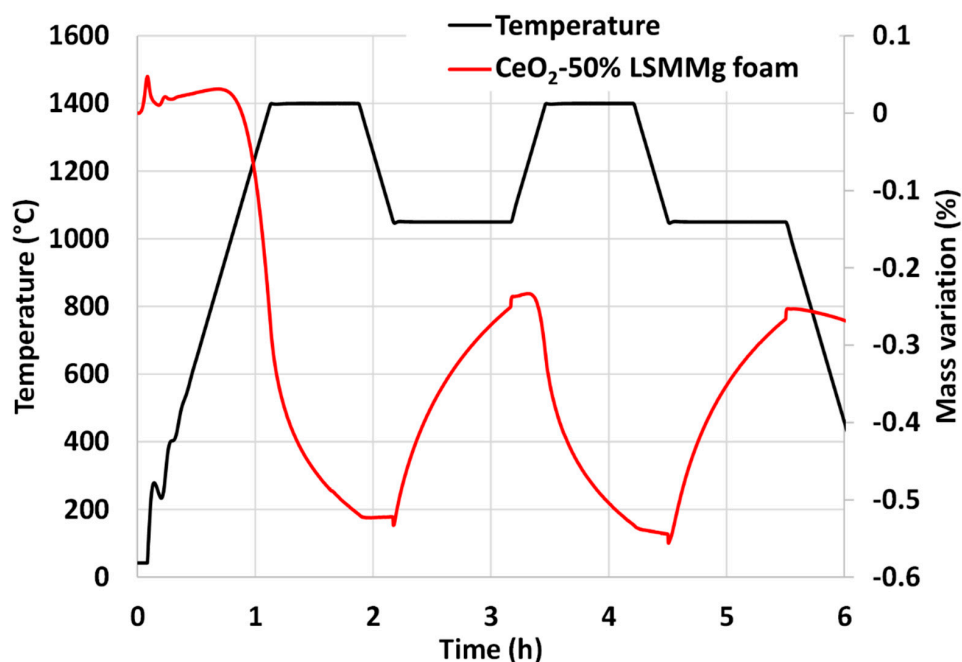
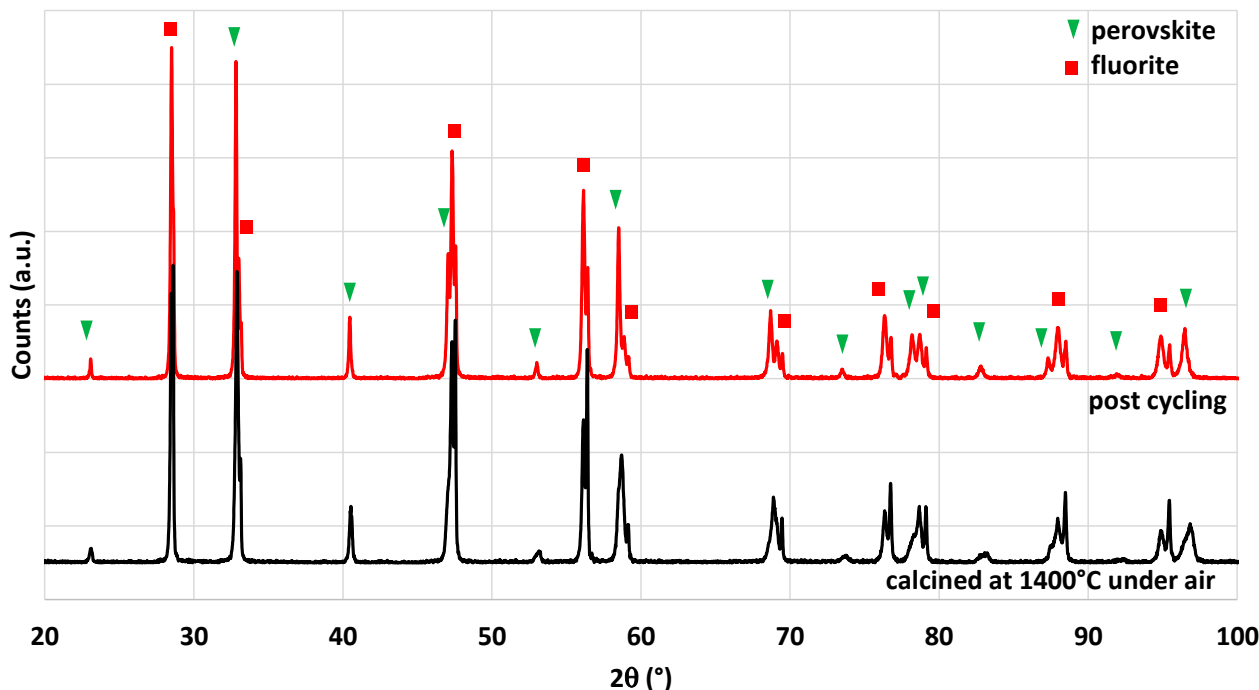


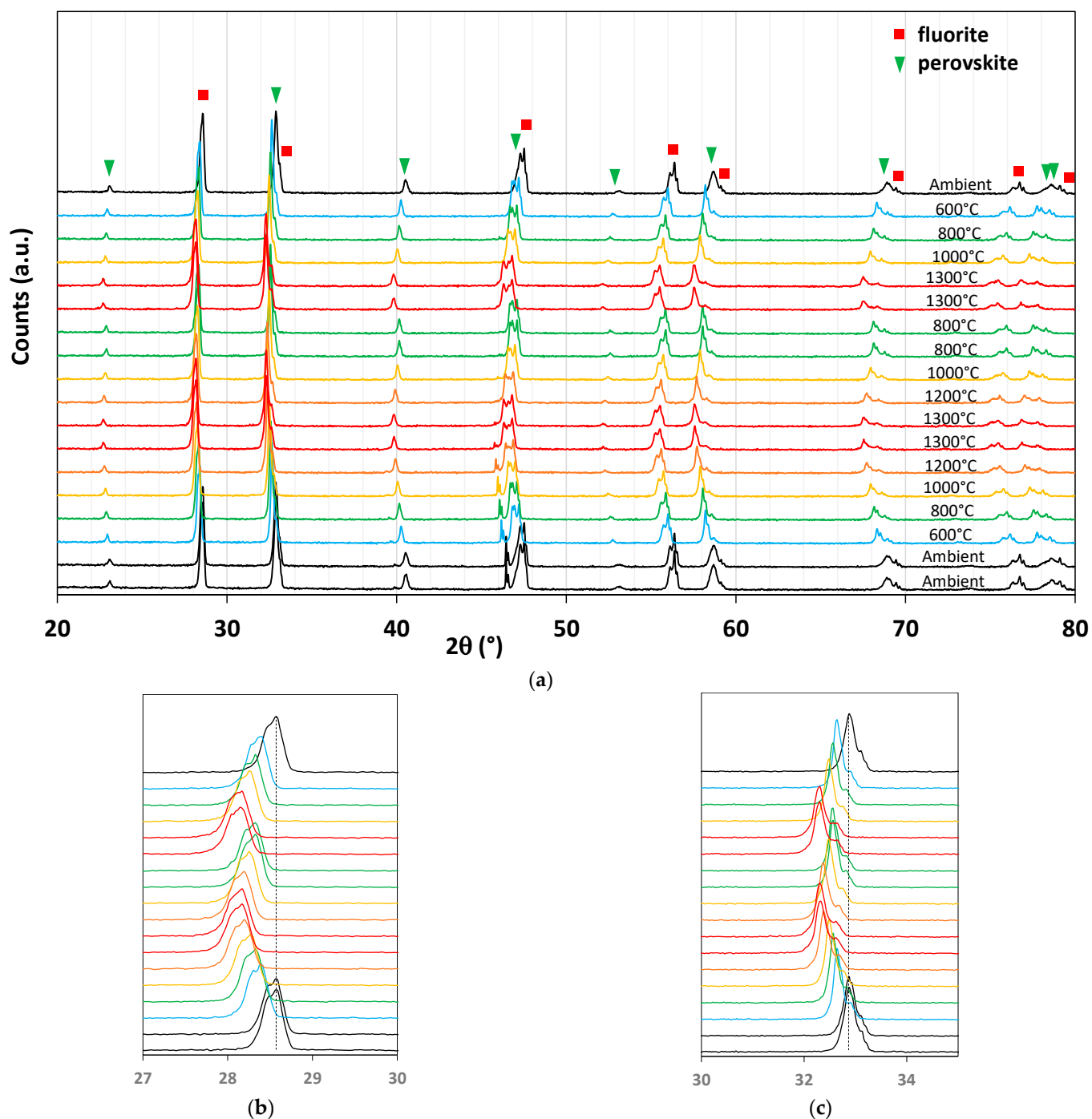
Figure 17. TGA of  $\text{CeO}_2/\text{LSMMg}$  (50/50 wt%) dual-phase foam during two consecutive  $\text{CO}_2$ -splitting cycles.

XRD patterns of the crushed foam calcined at 1400 °C and after two thermochemical cycles are compared in Figure 18. The two phases corresponding to CeO<sub>2</sub> and LSMMg are observed without any by-product before and after the thermochemical cycles, which confirms the stability of the material and the absence of reaction between phases.



**Figure 18.** XRD patterns of CeO<sub>2</sub>/LSMMg (50/50 wt%) dual-phase foam after calcination at 1400 °C and after two consecutive thermochemical CO<sub>2</sub>-splitting cycles.

In situ temperature-programmed XRD analysis during two thermochemical cycles was then performed. Figure 19a shows the superimposition of the diffraction patterns of the ground foam for the different temperatures. From bottom to top, the material was heated from room temperature to 1300 °C under nitrogen, then the temperature decreased to 800 °C, at which CO<sub>2</sub> was injected into the analysis chamber, followed by a similar second cycle. Diffraction peaks are shifted to low angles during reduction steps and to high angles during cooling and CO<sub>2</sub> exposure. These peak shifts correspond to the reduction and oxidation of the two materials. The beneficial synergistic effect described in previous work on La<sub>0.65</sub>Sr<sub>0.35</sub>MnO<sub>3</sub>–CeO<sub>2</sub> composite [43] was not observed in this study. The O<sub>2</sub> release and CO production yields are lower than those of pure LSMMg, and the effect of the oxygen diffusion channel by ceria was not observed. Indeed, Bork et al. [43] reported a shift of the LSM XRD peak to a lower angle upon cooling after the reduction step (corresponding to LSM reduction with oxygen transfer inducing simultaneous re-oxidation of ceria). They concluded that the ceria only played the role of oxygen diffusion channel, but this was not the case in this study as both phases were redox-active (Figure 19b,c).



**Figure 19.** XRD analysis of  $\text{CeO}_2/\text{LSMMg}$  (50/50 wt%) dual-phase foam during two consecutive  $\text{CO}_2$ -splitting cycles (a), zoom on a ceria peak (b), and zoom on a perovskite peak (c).

In summary, to conclude this section on the  $\text{CeO}_2/\text{LSMMg}$  (50/50) composite material, no synergistic effect was observed and CO production was not enhanced compared to the pure LSMMg phase. However, fuel production performance was improved over pure ceria despite slower oxidation rates.

### 3. Materials and Methods

Several synthesis methods were employed to obtain the different oxides investigated in this study. Different methods were chosen mainly to compare the redox performance with the materials synthesized in other studies based on the same methods.

Coprecipitation of hydroxides has been used for the synthesis of nickel ferrite and  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  oxides. Nitrates precursors ( $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ , purity 99.5%,  $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , purity 99.5%, and  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ , purity  $\geq 98\%$  from Sigma-Aldrich) were weighted to adapt to the desired stoichiometric content and dissolved in distilled water by stirring. Polyethylene glycol (purity  $\geq 99\%$ ) was added to the solution under stirring to improve the final powder morphology. Sodium hydroxide (NaOH) was then poured into the solution until the pH value reached 11. The metallic cations were hydrolyzed and formed a precipitate of metal hydroxides. The precipitate was filtered and washed several times until the pH of the rinsing solution decreased to 7. The precipitate was then dried in an oven at  $80\text{ }^\circ\text{C}$  overnight and calcined at  $1000\text{ }^\circ\text{C}$  for 1 h. The solid was ground in a mortar and calcined at  $1400\text{ }^\circ\text{C}$  (or  $1300\text{ }^\circ\text{C}$  for  $\text{Ce}_{0.9}\text{Fe}_{0.1}\text{O}_2$  to alleviate sintering) to form the final oxide.

The second synthesis method employed was the modified Pechini synthesis described elsewhere [48]. Briefly, metallic nitrate precursors were dissolved in distilled water by stirring for 10 min. The citric acid (purity  $\geq 99.5\%$ ) was then added to the solution under stirring with a molar concentration twice the cations concentration ( $n_{\text{CA}} = 2 \times n_{\text{cations}}$ ). Ethylene glycol (purity  $\geq 99\%$ ) was then poured into the solution with a molar concentration four times the cations concentration ( $n_{\text{EG}} = 4 \times n_{\text{cations}}$ ). The solution was stirred and heated up to  $90\text{ }^\circ\text{C}$  until gelation (brownish color). The gel was then heated in a furnace at  $250\text{ }^\circ\text{C}$  for 2 h. The resulting powder was crushed in an agate mortar and calcined in air at  $1000\text{ }^\circ\text{C}$  for 1 h. To avoid any sintering effect during the thermal analysis, the oxides were then calcined and stabilized in air for 1 h at either  $1300\text{ }^\circ\text{C}$  or  $1400\text{ }^\circ\text{C}$  corresponding to the maximum temperature of the thermogravimetric analysis (TGA).

Another synthesis technique employed was the evaporation-to-dryness method used to synthesize  $\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$  [37]. The metal nitrate precursors were dissolved in distilled water by stirring. The solution was then heated until complete evaporation of the water and the solid residue was first calcined at  $1100\text{ }^\circ\text{C}$  for 2 h, then at  $1400\text{ }^\circ\text{C}$  for 10 h.

The last method employed was a solid-state synthesis using a vibrating milling apparatus (mixer mill MM400 Retsch). To synthesize  $\text{Ce}_{0.2}\text{Sr}_{1.8}\text{MnO}_4$ , the selected precursors were  $\text{CeO}_2$  commercial powder (99.9%,  $<5\text{ }\mu\text{m}$ , Sigma-Aldrich),  $\text{MnO}$  ( $\geq 99\%$ ) commercial powder, and  $\text{SrCO}_3$  ( $\geq 99.9\%$ ). The precursors were weighed and placed in the mixer mill with water (37% wt). The paste was milled for 30 min at a frequency of 30 Hz. The paste obtained was dried at  $110\text{ }^\circ\text{C}$  for 6 h and calcined at  $1400\text{ }^\circ\text{C}$  for 3 h.

The reticulated foams were obtained from the corresponding powders by the replication method already described elsewhere [49]. The oxide powders were ground together in the case of dual-phase materials in a mortar and used in the slurry formulation. The different oxides were weighed and poured into distilled water together with the porogen (carbon fibers Sigrafil<sup>®</sup> M80), the dispersant (Dolapix<sup>®</sup>), and the binder (polyvinyl alcohol, PVA, Mowiol<sup>®</sup> 4-88). The aqueous slurry was then used to coat/impregnate a sacrificial polyurethane (PU, Bulpren) foam (10 ppi, pore per inch) serving as a template. The PU foam was immersed in the slurry and dried overnight. The as-impregnated foam was then slowly heated up to  $1000\text{ }^\circ\text{C}$  (at  $2.5\text{ }^\circ\text{C}/\text{min}$ ) to remove all the organic components from the material and annealed at  $1400\text{ }^\circ\text{C}$  for 3 h to achieve sintering which provides mechanical resistance to the ceramic foam.

The structure and microstructure of the prepared materials were fully characterized with the following series of techniques.

The crystalline structure was studied by X-ray diffraction (XRD) using a Panalytical X'PERT PRO diffractometer with the  $\text{Cu K}\alpha$  radiation ( $\lambda_{\text{Cu}} = 0.15406\text{ nm}$ , angular range =  $20\text{--}80^\circ$ ,  $2\theta$ , tube current 20 mA, potential 40 kV).

High-temperature X-ray diffractometer (HT XRD—Empyrean Panalytical with Anton-Paar HTK1 chamber) equipped with  $\text{Cu K}\alpha$  radiation was used for in situ investigation of phase reactivity/transitions during redox cycling (tube current 20 mA, potential 45 kV, angular range =  $20\text{--}80^\circ$ ,  $2\theta$ , angle variation  $0.02626^\circ/\text{s}$ ). The powder was heated up to

1300 or 1400 °C under N<sub>2</sub> flow for the reduction step and cooled down to either 1000 °C or 800 °C for the oxidation step with CO<sub>2</sub> injection.

The morphology of the materials was observed with a Field Emission Scanning Electron Microscope (FESEM–Hitachi S4800) used to examine the microstructure of foams.

A chemical elementary cartography was realized by EDS analysis (Energy Dispersive X-ray Spectroscopy) to estimate the chemical composition and to observe cations distribution (surface mapping) in the material. Analyses were performed using a Zeiss Sigma 300. The accelerating voltage used was 15 keV.

To complete the materials characterization, their thermochemical activity during two-step CO<sub>2</sub>-splitting cycles was investigated. Thermogravimetric analysis (TGA, SETARAM Setsys Evo 1750) was used to measure the mass variations (amount of oxygen exchanged) associated with the reduction and oxidation steps of thermochemical cycles. About 100 mg of material (exact mass determined with a high precision balance) was introduced in a platinum crucible hung to the microbalance with platinum suspensions and placed inside the furnace chamber. Residual air in the furnace chamber was eliminated by preliminary pumping to operate in an inert atmosphere during the reduction step with a low oxygen partial pressure ( $p_{O_2} \sim 10^{-5}$  bar). The sample was then heated under a flow of argon (99.999% purity, <2 ppm of O<sub>2</sub>, 20 mL/min) with a heating rate of 20 °C/min up to the selected set-point temperature, and the variation in mass was recorded continuously. The successive steps of the cycles were carried out at different temperatures (typically 1400 °C or 1300 °C dwelled for 45 min during the reduction step and 1050 °C dwelled for 1 h during the CO<sub>2</sub>-splitting step), and CO<sub>2</sub> (99.995% purity) was injected during the oxidation step (with a mole fraction of 50% CO<sub>2</sub> in argon and with a CO<sub>2</sub> flowrate of 10 mL/min). The same cycling conditions were applied for all materials for comparison purposes. The maximum cycle temperature was a good compromise because it was high enough to guarantee sufficient reduction yield while maintaining the thermal stability of the materials. Blank run correction (using an empty crucible) was performed to correct the small baseline mass drift during non-isothermal heating caused by thermal gas expansion and buoyancy effects on the crucible.

The mass variations associated with the temperature-programmed cycles correspond to the oxygen release during reduction (in the form of O<sub>2</sub>) and the oxygen uptake (in the form of O atoms) during CO<sub>2</sub>-splitting (equivalent to the CO produced), which allows calculating both the reduction yield and the CO production yield. Accordingly, the O<sub>2</sub> and CO production yields (expressed in mol/g<sub>oxide</sub>) are calculated as follows:

$$n_{O_2} = \Delta m_{red} / (M_{O_2} \cdot m_{oxide}) \quad (3)$$

$$n_{CO} = \Delta m_{ox} / (M_O \cdot m_{oxide}) \quad (4)$$

where  $\Delta m$  is the sample mass variation measured during the reduction or the oxidation step (g),  $M$  is the molecular weight of O<sub>2</sub> (32 g/mol) or O (16 g/mol), and  $m_{oxide}$  is the sample mass loaded in the crucible (g).

#### 4. Conclusions

This study investigated different classes of redox-active compounds based on spinel ferrite, fluorite, and perovskite oxide structures for the generation of synthetic solar fuels via two-step thermochemical cycles. The process ultimately aims to use concentrated solar energy as a high-temperature process heating source for thermochemical reactions. First, state-of-the-art redox materials consisting of Ni-ferrite and ceria were shaped as reticulated foams to compare their thermochemical performance during cycling with a reduction step temperature of 1400 °C and a re-oxidation step under CO<sub>2</sub> (50% mole fraction) at 1050 °C. As expected, the NiFe<sub>2</sub>O<sub>4</sub> foam produced more CO than its ceria counterpart, at the expense of slower kinetics during the oxidation step. Other single-phase oxides were then considered (including Ce<sub>0.9</sub>Fe<sub>0.1</sub>O<sub>2</sub>, Ca<sub>0.5</sub>Ce<sub>0.5</sub>MnO<sub>3</sub>, Ce<sub>0.2</sub>Sr<sub>1.8</sub>MnO<sub>4</sub>, and Sm<sub>0.6</sub>Ca<sub>0.4</sub>Mn<sub>0.8</sub>Al<sub>0.2</sub>O<sub>3</sub>) as they were previously identified in separate studies as highly

reactive materials for CO<sub>2</sub> splitting, with reported performance significantly superior to that of the benchmark ceria or lanthanum manganite perovskites. Nevertheless, these results were not confirmed in this work as the materials showed limited fuel production and lower overall redox performance in comparison with La<sub>0.5</sub>Sr<sub>0.5</sub>Mn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub>, due to sintering effects or kinetic limitations during oxidation which were never highlighted before.

In order to improve the performance of materials, dual-phase composites were finally considered, in particular ceria/ferrite and ceria/perovskite composites. The addition of NiFe<sub>2</sub>O<sub>4</sub> to ceria did not prove to be beneficial in improving redox performance. Among the perovskites considered (La<sub>0.5</sub>Sr<sub>0.5</sub>Mn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub>, Ca<sub>0.5</sub>Sr<sub>0.5</sub>MnO<sub>3</sub>, Y<sub>0.5</sub>Sr<sub>0.5</sub>MnO<sub>3</sub>), it was shown that the addition of 50 wt% of La<sub>0.5</sub>Sr<sub>0.5</sub>Mn<sub>0.9</sub>Mg<sub>0.1</sub>O<sub>3</sub> improved fuel yield at the expense of slower oxidation kinetics compared to pure ceria. Interestingly, the same composite formulation shaped as a reticulated foam with dual-scale porosity maintained its fuel production capacity (~180 μmol/g) without altering the reaction kinetics. However, no synergistic effect between the two phases was demonstrated.

The consideration of composite materials remains a promising option to improve the oxygen-exchange capacity of redox materials for the production of renewable fuels. When mixing redox-active phases, synergistic effects resulting from bulk-to-surface oxygen transfer in the composite materials are expected, by which the redox properties could be modified and ultimately the fuel productivity could be improved. The expected beneficial effects of dual-phase composites therefore deserve to be thoroughly assessed and further demonstrated.

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## References

1. Funk, J.; Reinstrom, R. Energy requirements in the production of hydrogen from water. *Ind. Eng. Chem. Process Des. Dev.* **1966**, *5*, 336–342. [[CrossRef](#)]
2. Han, S.; Kang, T.; Joo, O.; Jung, K. Water splitting for hydrogen production with ferrites. *Sol. Energy* **2007**, *81*, 623. [[CrossRef](#)]
3. Fresno, F.; Saaverda, R.; Gomez-Mancebo, M.; Vidal, A.; Sanchez, M.; Rucandio, I.; Quejido, A.; Romero, M. Solar hydrogen production by two-step thermochemical cycles: Evaluation of the activity of commercial ferrites. *Int. J. Hydrogen Energy* **2009**, *23*, 1185–1191. [[CrossRef](#)]
4. Kodama, T.; Gokon, N.; Yamamoto, R. Thermochemical two-step water splitting by ZrO<sub>2</sub>-supported Ni<sub>x</sub>Fe<sub>3-x</sub>O<sub>4</sub> for solar hydrogen production. *Sol. Energy* **2008**, *82*, 73–79. [[CrossRef](#)]
5. Jung, Y.; Oh, S.; Han, G.; Chen, K.; Seo, T. Dish type solar thermal system for hydrogen production with two-step water-splitting reaction. In Proceedings of the Solar Paces Conference, Perpignan, France, 21–24 September 2010; Volume 179.



6. Gokon, N.; Kondo, N.; Mataga, T.; Kodama, T. Internally circulating fluidized bed reactor with NiFe<sub>2</sub>O<sub>4</sub> particles for thermochemical water-splitting. In Proceedings of the Solar Paces Conference, Berlin, Germany, 15–18 September 2009; pp. 46–56.
7. Abanades, S.; Flamant, G. Thermochemical hydrogen production from a two-step solar-driven water-splitting cycle based on cerium oxides. *Sol. Energy* **2006**, *80*, 1611–1623. [[CrossRef](#)]
8. Marxer, D.; Furler, P.; Takacs, M.; Steinfeld, A. Solar thermochemical splitting of CO<sub>2</sub> into separate streams of CO and O<sub>2</sub> with high selectivity, stability, conversion, and efficiency. *Energy Environ. Sci.* **2017**, *10*, 1142–1149. [[CrossRef](#)]
9. Schappi, R.; Rutz, D.; Dahler, F.; Muroyama, A.; Haueter, P.; Lilliestam, J.; Patt, A.; Furler, P.; Steinfeld, A. Drop-in fuels from sunlight and air. *Nature* **2022**, *601*, 63–68. [[CrossRef](#)]
10. Chueh, W.; Falter, C.; Abbott, M.; Scipio, D.; Furler, P.; Haile, S. High flux solar driven thermochemical dissociation of CO<sub>2</sub> and H<sub>2</sub>O using non-stoichiometric ceria. *Science* **2010**, *330*, 1797–1801. [[CrossRef](#)]
11. Le Gal, A.; Abanades, S. Catalytic investigation of ceria-zirconia solid solutions for solar hydrogen production. *Int. J. Hydrogen Energy* **2011**, *36*, 4739–4748. [[CrossRef](#)]
12. Arifin, D.; Ambrosini, A.; Wilson, S.; Mandal, B.; Muhich, C.; Weimer, A. Investigation of Zr, Gd/Zr, and Pr/Zr-doped ceria for the redox splitting of water. *Int. J. Hydrogen Energy* **2020**, *45*, 160–174. [[CrossRef](#)]
13. Pullar, R.; Novais, R.; Caetano, A.; Barreiros, M.; Abanades, S.; Costa Oliveira, F. A review of solar thermochemical CO<sub>2</sub> splitting using ceria-based ceramics with designed morphologies and microstructures. *Front. Chem.* **2019**, *7*, 601. [[CrossRef](#)] [[PubMed](#)]
14. Furler, P.; Scheffe, J.; Steinfeld, A. Syngas production by simultaneous splitting of H<sub>2</sub>O and CO<sub>2</sub> via ceria redox reaction in a high-temperature solar reactor powered by a high flux solar simulator. *Energy Environ. Sci.* **2012**, *5*, 6098–6103. [[CrossRef](#)]
15. Furler, P.; Scheffe, J.; Gorbar, M.; Moes, L.; Vogt, U.; Steinfeld, A. Solar thermochemical CO<sub>2</sub> splitting utilizing a reticulated porous ceria redox system. *Energy Fuels* **2012**, *26*, 7051–7059. [[CrossRef](#)]
16. Haeussler, A.; Abanades, S.; Julbe, A.; Jouannaux, J.; Drobek, M.; Ayrat, A.; Cartoixa, B. Remarkable Performance of Microstructured Ceria Foams for Thermochemical Splitting of H<sub>2</sub>O and CO<sub>2</sub> in a Novel High-Temperature Solar Reactor. *Chem. Eng. Res. Des.* **2020**, *156*, 311–323. [[CrossRef](#)]
17. Haeussler, A.; Abanades, S.; Costa Oliveira, F.A.; Barreiros, M.A.; Caetano, A.P.F.; Novais, R.M.; Pullar, R.C. Solar Redox Cycling of Ceria Structures Based on Fiber Boards, Foams, and Biomimetic Cork-Derived Ecoceramics for Two-Step Thermochemical H<sub>2</sub>O and CO<sub>2</sub> Splitting. *Energy Fuels* **2020**, *34*, 9037–9049. [[CrossRef](#)]
18. Venstrom, L.J.; Petkovich, N.; Rudisill, S.; Stein, A.; Davidson, J.H. The effects of morphology on the oxidation of ceria by water and carbon dioxide. *J. Sol. Energy Eng.* **2012**, *134*, 011005. [[CrossRef](#)]
19. Abanades, S.; Haeussler, A. Two-Step Thermochemical Cycles Using Fibrous Ceria Pellets for H<sub>2</sub> Production and CO<sub>2</sub> Reduction in Packed-Bed Solar Reactors. *Sustain. Mater. Technol.* **2021**, *29*, e00328. [[CrossRef](#)]
20. Gibbons, W.; Venstrom, L.; De Smith, R.; Davidson, J.; Jackson, G. Ceria-based electrospun fibers for renewable fuel production via two-step thermal redox cycles for carbon dioxide splitting. *Phys. Chem. Chem. Phys.* **2014**, *16*, 14271–14280. [[CrossRef](#)]
21. Abanades, S.; Haeussler, A.; Julbe, A. Synthesis and thermochemical redox cycling of porous ceria microspheres for renewable fuels production from solar-aided water-splitting and CO<sub>2</sub> utilization. *Appl. Phys. Lett.* **2021**, *119*, 023902. [[CrossRef](#)]
22. Takalkar, G.; Bhosale, R.; AlMomani, F.; Rashid, S.; Qiblawey, H.; Saad, M.; Khraisheh, M.; Kumar, G.; Gupta, R.; Shende, R. Thermochemical splitting of CO<sub>2</sub> using solution combustion synthesized lanthanum-strontium-manganese perovskites. *Fuel* **2021**, *285*, 119154. [[CrossRef](#)]
23. Haeussler, A.; Julbe, A.; Abanades, S. Investigation of reactive perovskite materials for solar fuel production via two-step redox cycles: Thermochemical activity, thermodynamic properties and reduction kinetics. *Mater. Chem. Phys.* **2022**, *276*, 125358. [[CrossRef](#)]
24. Barcellos, D.; Sanders, M.; Tong, J.; McDaniel, A.; O’Hayre, R. BaCe<sub>0.25</sub>Mn<sub>0.75</sub>O<sub>3</sub>—A promising perovskite type oxide for solar thermochemical hydrogen production. *Energy Environ. Sci.* **2018**, *11*, 3256–3265. [[CrossRef](#)]
25. Zhang, D.; De Santiago, H.; Xu, B.; Liu, C.; Trindell, J.; Li, W.; Park, J.; Rodriguez, M.; Coker, E.; Sugar, J.; et al. Compositionally complex perovskite oxides for solar thermochemical water splitting. *Chem. Mater.* **2023**, *35*, 1901–1915. [[CrossRef](#)]
26. Mao, Y.; Gao, Y.; Dong, W.; Wu, H.; Song, Z.; Zhao, X.; Sun, J.; Wang, W. Hydrogen production via two-step water splitting thermochemical cycle based on metal oxide—A review. *Appl. Energy* **2020**, *267*, 114860. [[CrossRef](#)]
27. Abanades, S. Metal oxides applied to thermochemical water-splitting for hydrogen production using concentrated solar energy. *Chem. Eng.* **2019**, *3*, 63. [[CrossRef](#)]
28. Abanades, S. Redox Cycles, Active Materials, and Reactors Applied to Water and Carbon Dioxide Splitting for Solar Thermochemical Fuel Production: A Review. *Energies* **2022**, *15*, 7061. [[CrossRef](#)]
29. Gokon, N.; Murayama, H.; Nagasaki, A.; Kodama, T. Thermochemical two-step water-splitting cycles by monoclinic ZrO<sub>2</sub>-supported NiFe<sub>2</sub>O<sub>4</sub> and Fe<sub>3</sub>O<sub>4</sub> powders and ceramic foam devices. *Sol. Energy* **2009**, *89*, 527–537. [[CrossRef](#)]
30. Gokon, N.; Takahashi, S.; Yamamoto, H.; Kodama, T. Thermochemical two-step water-splitting reactor with internally circulating fluidized bed for thermal reduction of ferrites particles. *Int. J. Hydrogen Energy* **2008**, *33*, 2189–2199. [[CrossRef](#)]
31. Allendorf, M.; Diver, R.; Siegel, N.; Miller, J. Two-step water splitting using mixed metal ferrites: Thermodynamic analysis and characterization of synthesized materials. *Energy Fuels* **2008**, *22*, 4115–4124. [[CrossRef](#)]
32. Lorentzou, S.; Zygogianni, A.; Pagkoura, C. Hydrosol-plant: Structured redox reactors for H<sub>2</sub> production from solar thermochemical H<sub>2</sub>O splitting. *AIP Conf. Proc.* **2018**, *2033*, 130010. [[CrossRef](#)]

33. Orfila, M.; Sanz, D.; Linares, M.; Molina, R.; Sanz, R.; Marugan, J.; Botas, J.A. H<sub>2</sub> production by thermochemical water splitting with reticulated porous structures of ceria-based mixed oxide materials. *Int. J. Hydrogen Energy* **2021**, *46*, 17458–17471. [[CrossRef](#)]
34. Al-Taweel, S.; Nadeem, M.; Idriss, H. A study of Ce<sub>x</sub>Fe<sub>1-x</sub>O<sub>2</sub> as a reducible oxide for the thermal hydrogen production from water. *Energy Technol.* **2022**, *10*, 2100491. [[CrossRef](#)]
35. Kümmerle, E.; Heger, G. The structures of C-Ce<sub>2</sub>O<sub>3+δ</sub>, Ce<sub>7</sub>O<sub>12</sub>, and Ce<sub>11</sub>O<sub>20</sub>. *J. Solid-State Chem.* **1999**, *147*, 485–500. [[CrossRef](#)]
36. Jouannaux, J.; Haeussler, A.; Drobek, M.; Ayrat, A.; Abanades, S.; Julbe, A. Lanthanum manganite perovskite ceramic powders for CO<sub>2</sub> splitting: Influence of Pechini synthesis parameters on sinterability and reactivity. *Ceram. Int.* **2019**, *45*, 15636–45648. [[CrossRef](#)]
37. Barcellos, D.; Coury, F.; Emery, A.; Sanders, M.; Tong, J.; McDaniel, A.; Wolverton, C.; Kaufman, M.; O’Hayre, R. Phase identification of the layered perovskite Ce<sub>x</sub>Sr<sub>2-x</sub>MnO<sub>4</sub> and application for solar thermochemical water splitting. *Inorg. Chem.* **2019**, *58*, 7705–7714. [[CrossRef](#)] [[PubMed](#)]
38. Bergeson-Keller, A.; Sanders, M.; O’Hayre, R. Reduction thermodynamics of Sr<sub>1-x</sub>Ce<sub>x</sub>MnO<sub>3</sub> and Ce<sub>x</sub>Sr<sub>2-x</sub>MnO<sub>4</sub> perovskites for solar thermochemical hydrogen production. *Energy Technol.* **2022**, *10*, 2100515. [[CrossRef](#)]
39. Bagherisereshki, E.; Tran, J.; Lei, F.; Au Yeung, N. Investigation into SrO/SrCO<sub>3</sub> for high temperature thermochemical energy storage. *Sol. Energy* **2018**, *160*, 85–93. [[CrossRef](#)]
40. Gautam, G.; Stetchel, E.; Carter, E. Exploring Ca-Ce-M-O (M=3d transition metal) oxide perovskites for solar thermochemical applications. *Chem. Mater.* **2020**, *32*, 9964–9982. [[CrossRef](#)]
41. Dukic, J.; Boscovic, S.; Matovic, B. Crystal structure of Ce-doped CaMnO<sub>3</sub> perovskite. *Ceram. Int.* **2009**, *35*, 787–790. [[CrossRef](#)]
42. Gao, K.; Liu, X.; Jiang, Z.; Zheng, H.; Song, C.; Wang, X.; Tian, C.; Dang, C.; Sun, N.; Xuan, Y. Direct solar thermochemical CO<sub>2</sub> splitting based on Ca- and Al-doped SmMnO<sub>3</sub> perovskites: Ultrahigh CO yield within small temperature swing. *Renew. Energy* **2022**, *194*, 482–494. [[CrossRef](#)]
43. Bork, A.; Carrillo, A.; Hood, Z.; Yildiz, B.; Rupp, J. Oxygen exchange in dual-phase La<sub>0.66</sub>Sr<sub>0.35</sub>MnO<sub>3</sub>-CeO<sub>2</sub> composites for solar thermochemical fuel production. *Appl. Mater. Interfaces* **2020**, *12*, 32622–32632. [[CrossRef](#)] [[PubMed](#)]
44. Huang, X.; Ni, C.; Zhao, G.; Irvine, J. Oxygen storage capacity and thermal stability of the CuMnO<sub>2</sub>-CeO<sub>2</sub> composite system. *J. Mater. Chem. A* **2015**, *3*, 12958–12964. [[CrossRef](#)]
45. Zhang, D.; Kawada, T.; Yoshioka, F.; Machida, M. Oxygen gateway effect of CeO<sub>2</sub>/La<sub>2</sub>O<sub>2</sub>SO<sub>4</sub> composite oxygen storage materials. *ACS Omega* **2016**, *1*, 789–798. [[CrossRef](#)]
46. Machida, M.; Kawada, T.; Fujii, H.; Hinokuma, S. The role of CeO<sub>2</sub> as a gateway for oxygen storage over CeO<sub>2</sub> grafted Fe<sub>2</sub>O<sub>3</sub> composite materials. *J. Phys. Chem. C* **2015**, *119*, 24932–24941. [[CrossRef](#)]
47. Haeussler, A.; Abanades, S.; Julbe, A.; Jouannaux, J.; Cartoixa, B. Two-step CO<sub>2</sub> and H<sub>2</sub>O splitting using perovskite coated ceria foam for enhanced green fuel production in a porous volumetric solar reactor. *J. CO<sub>2</sub> Util.* **2020**, *41*, 101257. [[CrossRef](#)]
48. Sunde, T.; Grande, T.; Einarsrud, M. Modified Pechini synthesis of oxide powders and thin films. In *Handbook of Sol-Gel Science and Technology*; Klein, L., Aparicio, M., Jitianu, A., Eds.; Springer International Publishing: Cham, Switzerland, 2018; pp. 1089–1118. [[CrossRef](#)]
49. Schwartzwalder, K.; Somers, H.; Somers, V. Method of Making Porous Ceramic Articles. U.S. Patent 3090094A, 21 May 1963.

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