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Galvanic-Replacement-Induced Introduction of an Hetero-ligand into Bimetallic and Trimetallic Nanoclusters†

Jian-Hong Liao,^a Tzu-Hao Chiu,^a Hao Liang,^b Samia Kahlal,^b Jean-Yves Saillard ^{*b} and C. W. Liu ^{*a}

Heteroleptic eight-electron silver-rich alloy nanoclusters, [Au@Au₄Ag₁₂(dtp)₇(PPh₃)₄]²⁺ (1), and [Pt@Au₄Ag₁₁(dtp)₇(PPh₃)₄] (2) were successfully synthesized via galvanic-replacement reaction of 1,1-dithiolate-protected M@Ag₂₀ (M = Au, Pt) nanoclusters with Au(I)-phosphine salts leading to the alteration of the cluster nuclearity and geometry of shell skeletons but retaining the same 8-electron count.

Ligand-protected metal nanoclusters (NCs) with atom-precise compositions have undergone many developments in their fine-tuning optical, catalytic, and biomedical applications.¹⁻³ These NCs exhibit optical properties that are correlated to several factors, such as the geometry and electron count of their kernel, the presence of a dopant (if any), the nature of the ligand-protected layer, and the possible supramolecular assembly of the NCs.⁴⁻¹⁰ Thus, molecular engineering on these features has become a fascinating task. A commonly used method for alloying homo-metallic NCs is to introduce hetero-metals into their framework via galvanic or anti-galvanic exchanges.¹¹⁻¹² The yielded alloys generally maintain the original kernel skeletons, and only one or a few positions are exchanged, yielding single-site or multiple-site replacements.¹³⁻²² For example, the silver atom located in the middle of the central Ag₁₂ icosahedron of [Ag₂₅(SR)₁₈]⁻ can be replaced by gold, to form [AuAg₂₄(SR)₁₈]⁻, leading to a significant enhancement of photoluminescence.¹³ Moreover, multiple-site replacements on the Ag₁₂ icosahedron can also occur upon adding extra Au atoms to [AuAg₂₀{Se₂P(OⁱPr)₂}]₁₂⁺, forming [Au₃Ag₁₇{Se₂P(OⁱPr)₂}]₁₂⁺ accompanied with symmetry changes.¹⁷ A similar idea was used to dope Ag₂₉(SSR)₁₂(PPh₃)₄ stepwise, first with copper to produce Ag₁₇Cu₁₂(SSR)₁₂(PPh₃)₄, and then with gold to reach the trimetallic alloy

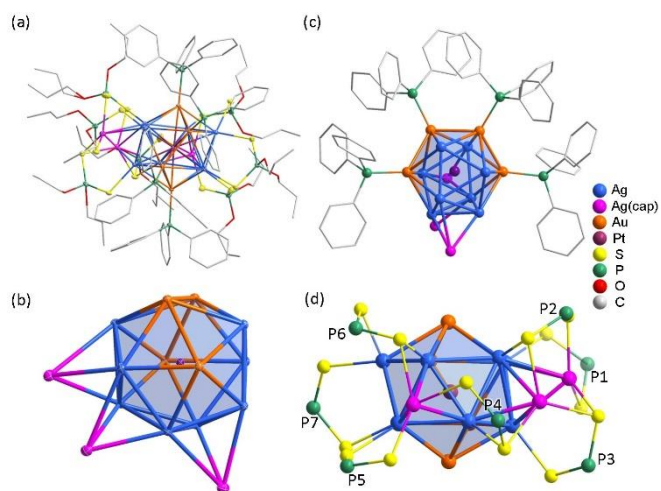
AuAg₁₆Cu₁₂(SSR)₁₂(PPh₃)₄ (SSR = benzene-1, 3-dithiolate).²³ The three isoelectronic NCs retain the same structural framework.

Thus, stepwise doping is a successful strategy to prepare trimetallic NCs.^{20-21,23-24} Following the same strategy, we have isolated bimetallic and trimetallic M₂₀ alloys, [Cu_{3.5}Ag_{16.5}(dtp)₁₂] and [Cu_{2.5}AuAg_{16.5}(dtp)₁₂] (dtp = S₂P(OⁱPr)₂).²¹ It is worth noting that, in the above described examples, the metal nuclearity and the number of ligands are maintained upon doping.^{21,23} This can be possibly due to the innate bite angle of the dithiolate ligand, which provides less flexibility to the NC surface. In this context, increasing or decreasing the nuclearity of M₂₀ NCs is restricted to a certain extent, which may explain the fact that only homoleptic silver-rich M₂₀ or M₂₁ NCs could be isolated so far.^{16-18,20-21,25} Therefore, introducing hetero-donor ligands onto the cluster surface should be the key for allowing nuclearity modifications. In the broadly studied phosphine-protected Au NCs, the labile P-Au bond provides tremendous advantages in ligand exchange with other organic ligands such as thiolate ligands, encouraging the development of heteroleptic NCs.²⁶ In contrast to the relatively weak Au-P bonding, the strong Au-S bonding leads to the formation of highly stable Au NCs, making them difficult to undergo phosphine-ligand exchange reactions. A recent study reports the successful introduction of diphosphine ligands onto the cluster surface by adding Au-diphosphine complexes to a thiolate-protected Au NC, [Au₂₃(SR)₁₆]⁻.²⁷ Motivated by this successful example, we adopted a similar strategy with Ag NCs. In this work, PPh₃ ligands have been introduced onto the NCs successfully together with the dopant. A bimetallic M₁₇ NC, [Au@Au₄Ag₁₂(dtp)₇(PPh₃)₄]²⁺ (1), and a trimetallic M₁₆ NC, [Pt@Au₄Ag₁₁(dtp)₇(PPh₃)₄] (2), have been isolated and characterized by X-ray crystallography, UV-Vis, and DFT calculations. Taking into account that Pt(0), Ag(0) and Au(0) are providing 0, 1 and 1 electron to the cluster count, respectively, and that the dtp ligands are formally anionic, both 1 and 2 are 8-electron NCs. To our knowledge, eight-

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compound **2a**.

Fig. 2 (a) The total structure of $[\text{Pt}@\text{Au}_4\text{Ag}_{11}\{\text{S}_2\text{P}(\text{O}^i\text{Pr})_2\}_7(\text{PPh}_3)_4]$, **2**. (b) The $\text{PtAu}_4\text{Ag}_{11}$ metal framework (Thermal ellipsoid with 50% probability). (c) With the four PPh_3 ligands (d) With the seven dtp ligands (propoxy groups omitted).

Its complete structure is shown in Fig. 2a. The M_{16} metal framework can be seen as a Pt-centered Au_4Ag_8 icosahedron capped by three silver atoms (Ag_{cap}). The average $\text{M}_{\text{cent}}\text{-M}_{\text{ico}}$ (2.7631(7) Å) and $\text{M}_{\text{ico}}\text{-M}_{\text{ico}}$ distances (2.9055(7) Å) in **2a** are similar to **1**, indicating the two icosahedral skeletons are almost identical. Unlike the four edge-bridging silvers presented in **1**, there are only three face-capping silvers in **2a** (Fig. 2b) with an average $\text{Ag}_{\text{cap}}\text{-M}_{\text{ico}}$ distance of 3.0183(9) Å. Although this does not affect the coordination mode of the four icosahedral Au atoms terminally bonded to PPh_3 ligands (average P-Au distance: 2.315(2) Å, Fig. 2c), it influences the orientation of the seven dtp ligands (Fig. 2d), which this time adopt $\eta^3(\mu_2, \mu_1)$ and $\eta^2(\mu_1, \mu_1)$ coordination mode for P2, P3, P4, P6 and P1, P5, P7, respectively. Therefore, both the M_{16} metal framework and the entire structure of **2** possess only C_1 symmetry.

From the perspective of the outer shell in **1** and **2**, the PPh_3 ligands coordinate radially to the icosahedron Au atoms, thus resulting in a linear P-Au- M_{cent} arrangement (the average angle

is 176.4° in **1** and 176.8° in **2**). It is of note that in the bicapped icosahedral metal framework of $[\text{MAu}_6\text{Ag}_8(\text{SAr}^F)_6(\text{PPh}_3)_6]$ ($\text{M} = \text{Pd}, \text{Pt}$, $\text{Ar}^F = 3,5\text{-(CF}_3)_2\text{C}_6\text{H}_3$), the six Au atoms also occupy icosahedral vertices in an Au_6 chair conformation, each being linearly coordinated to a PPh_3 ligand.³⁴ Thus, it is anticipated that linearly coordinated terminal ligands, such as alkyne, monothiol, or N-heterocyclic carbene could be introduced as hetero ligands in the galvanic or anti-galvanic replacement processes. The doping gold atoms occupy M_{ico} instead of M_{cap} positions in **1** and **2**, which fully agrees with a previous DFT investigation on 8-electron $[\text{MAG}_{20}(\text{dtp})_{12}]$ species.³⁵

Since in **1** and **2** the four PPh_3 ligands orient in the same way, *i.e.* that minimizing steric hindrance (Fig. 1c and 2c), the difference between the two structures lies in the fact that, with one bridging atom more than in the latter, the outer shell of the former is more crowded, resulting in edge-bridging Ag_{bri} rather than face-bridging Ag_{cap} atoms in the latter. In addition, it is also noted that the different formal valence charges of the superatomic kernels, $[\text{Au}@\text{Au}_4\text{Ag}_8]^{5+}$ in **1** and $[\text{Pt}@\text{Au}_4\text{Ag}_8]^{4+}$ in **2**, may have implications in their interaction with the outer shell. These results demonstrate that the number of capping metal atoms and the charge of the kernel significantly affect the ligand coordination when the heteroligand participates in the protective shell.

The ESI-MS of **2** (Fig. S2) shows two distributions that contain the molecular ion peaks of $[\text{PtAu}_4\text{Ag}_{11}(\text{dtp})_7(\text{PPh}_3)_4 + \text{Ag}^+]^+$ at m/z 4818.8408 (calc. m/z 4818.1578 for $[\mathbf{2a} + \text{Ag}^+]^+$) and $[\text{PtAu}_5\text{Ag}_{10}(\text{dtp})_7(\text{PPh}_3)_4 + \text{AgPPh}_3]^+$ at m/z 5170.2764 (calc. m/z 5170.3283 for $[\mathbf{2b} + \text{AgPPh}_3]^+$). Intriguingly these two peaks correspond to the compositions determined in the co-crystal of **2a** and **2b**. The other molecular ion peaks are likely to be species resulting from the fragmentation of the gas-phase-unstable NCs of **2**. The detailed assignment of each peak is given in Fig. S2.

The $^{31}\text{P}\{\text{H}\}$ NMR spectrum of **1** (Fig. S3) shows two resonances at around 102.5 and 37.2 ppm, associated with the dtp and PPh_3 ligands, respectively. In contrast, the $^{31}\text{P}\{\text{H}\}$ NMR spectrum of **2** reveals several chemical shifts from 103 to 109 ppm, attributable to the various chemical environments of dtp ligands in the less symmetrical structure **2**.

The UV-vis spectrum of **1** (Fig. 3a) shows a high-energy band at 263 nm corresponding to $\pi \rightarrow \pi^*$ transitions (LLCT). The

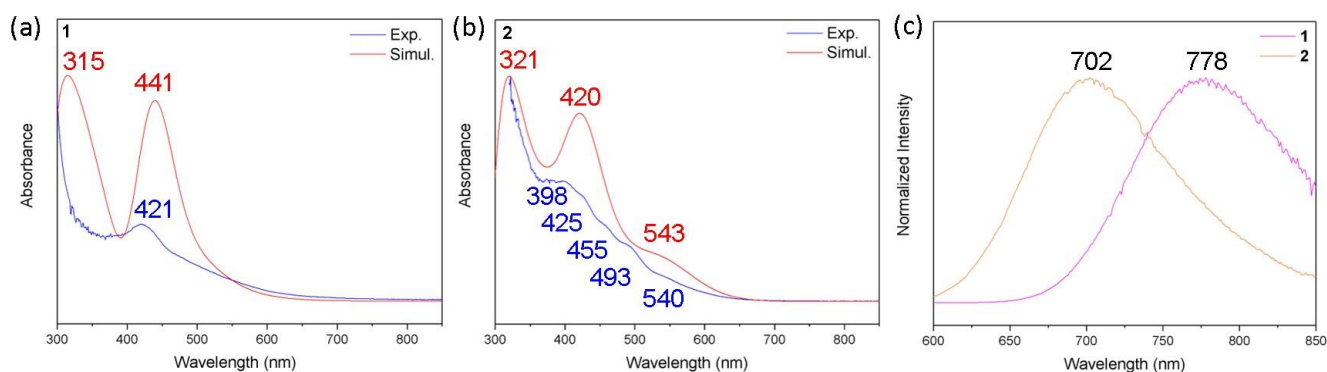


Fig. 3 (a) The absorption (blue) and simulated (red) spectra of **1**. (b) The absorption (blue) spectrum of **2** and the simulated (red) spectrum of **2a**. (c) The emission spectra of **1** (orange) and **2** (magenta) in 2-MeTHF glass at 77K.

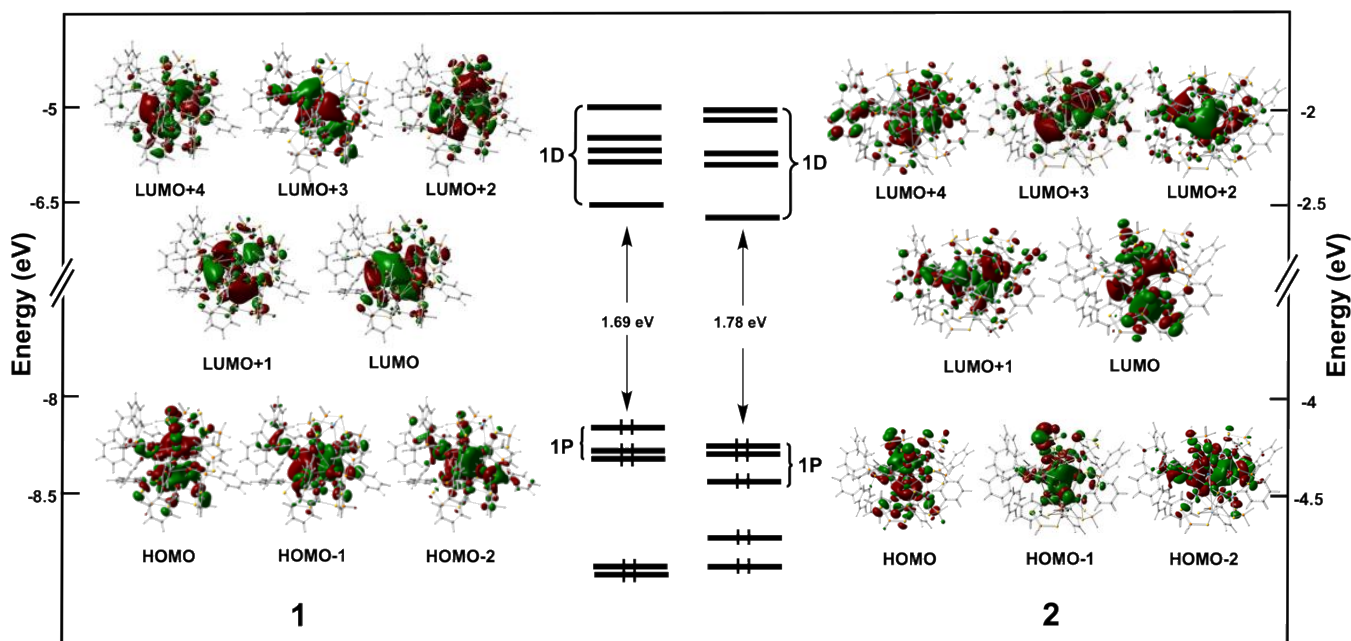


Fig. 4 Kohn-Sham frontier orbital diagrams of **1** and **2**. The low-lying 1S occupied orbitals are not shown.

low-energy absorption at 421 nm can be attributed to 1P \rightarrow 1D transitions (see below). The spectrum of **2** (Fig. 3b) displays five distinct absorption peaks at 398, 425, 455, 493, and 540 nm.

1 does not emit in solution at room temperature, but **2** reveals very weak emission at 725 nm. The glassy state of **1** and **2** at 77K shows red emission with maximal emission wavelengths of 778 and 702 nm, respectively (Fig. 3c). The lifetimes of **1** (Fig. S7) and **2** (Fig. S8) are 18 and 16 μ s, respectively, characteristic of phosphorescence. For comparison, another example of an eight-electron alloy, Au₄Ag₁₃(DPPM)₃(SR)₉, displayed red emission at \sim 695 nm with a short-lived lifetime of 3.31 ns in the solution state.³¹

Density functional theory (DFT) calculations were carried out at the BP86/Def2SVP level (see Computational Details section in the supporting information) to shed some light on the bonding and properties of compounds **1** and **2a**, in which the S₂P(OⁱPr)₂ ligands were replaced by S₂PH₂ in order to reduce computational cost. Selected computed data are given in Table 1. The optimized geometries of **1** and **2** are in good agreement with their X-ray counterparts. The M_{cent}-M_{ico} Wiberg bond index is typical for an 8-electron icosahedral system, whereas the M_{ico}-M_{bri/cap} one is indicative of very weak covalent interaction, leaving the outer M_{bri/cap} atoms as locally stable (nearly) planar tricoordinated 16-electron Ag(I) centers. The NAO charges of both NCs are similar, except for that of M_{cent}, as expected from their different nature (Au vs. Pt). The Kohn-Sham MO diagrams of **1** and **2** are shown in Fig. 4. They exhibit rather large HOMO-LUMO gaps, in line with their stability. In both NCs, the three highest occupied orbitals can be identified as the 1P shell, whereas the five lowest vacant ones correspond to the 1D manifold, thus confirming their *superatomic* 1S² 1P⁶ 1D⁰ configuration. TD-DFT calculations

performed at the B3LYP/Def2SVP level on **1** and **2a** allow for simulating satisfyingly the shapes of their UV-vis spectra (Fig. 3a and 3b). In the case of **1**, the band at 441 nm is of 1P \rightarrow 1D *superatomic* nature (MMCT), whereas that at 315 nm corresponds to MLCT transitions. In the case of **2**, both bands a 543 and 420 nm are of 1P \rightarrow 1D nature, and that at 321 nm is of MLCT character.

Conclusions

In summary, a new synthetic pathway for introducing hetero ligands into dithiolate-protected nanoclusters was investigated. Owing to the high stability of dithiolate-protected nanoclusters, it is not easy to directly add other organic ligands, such as phosphine ligands, to achieve substitutions. The galvanic replacement triggered the phosphine addition when the parent M@Ag₂₀ NCs react with the Au-phosphine complex, yielding bimetallic M₁₇ and trimetallic M₁₆ NCs, [Au@Au₄Ag₁₂(dtp)₇(PPh₃)₄]²⁺ and [Pt@Au₄Ag₁₁(dtp)₇(PPh₃)₄]. A prolonged reaction time in the same conditions does not yield a higher number of Au replacements. The results exhibit the modification of cluster surface, nuclearity, and also optical properties. The experimental absorption spectra are well-reproduced by TD-DFT calculations, where the low-energy bands at 441 nm (in **1**) and 420, 543 nm (in **2**) correspond to 1P \rightarrow 1D transitions. This method offers opportunities to design new heteroleptic NCs and maps out the possibility of NC surface engineering.

Conflicts of interest

There are no conflicts to declare.

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