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# Lifshitz phase: the microscopic structure of aqueous and ethanol mixtures of 1,n-diols

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## Abstract

We study by computer simulations binary mixtures of ethylene glycol and 1,3-propanediol with water or ethanol. Despite strong hydrogen bonding tendencies between all these molecules, we find that these mixtures are surprisingly homogeneous, in contrast with the strong micro-heterogeneity found in aqueous ethanol mixtures. The aqueous diol mixtures are found to be close to ideal mixing, with near-ideal Kirkwood-Buff integrals. Ethanol-diol mixtures show weak non-ideality. The origin of this unexpected randomness is due to the fact that the two hydrogen bonding hydroxyl groups of the 1,n-diol are tied by the neutral alkyl bond, which prevents the micro-segregation of the different types of hydroxyl groups. These findings suggest that random disorder can arise in presence of strong interactions -in contrast to the usual picture of random disorder as due to weak interactions between the components. They point to the important role of the molecular topology in tuning concentration fluctuations in complex liquids. We propose and justify herein the name of Lifshitz phases to designate such types of disordered systems.

## 1 Introduction

The concept of Lifshitz point was initially pointed out in a field theoretic context[1], as a triple point along the lambda-line separating a disordered phase from two

types of ordered phases. The peculiarity of this Lifshitz point is the emergence of a characteristic  $k \neq 0$  wave vector pre-peak in the scattering function  $I(k)$ , as one goes from the disordered phase, where the peak is at  $k = 0$ , to one of the ordered phases[1]. A particular property in the context of the Lifshitz point is that the transition between the disordered and the ordered phase is continuous, without the usual expected gap[1]. The emergence of a  $k \neq 0$  pre-peak in this particular context, is usually the sign of some sort of layered phase. This situation is often met in the context of micro-emulsions[2, 3], when going from the disordered phase to the lamellar phase[4]. This type of scenario is well captured by various lattice models[5] and field theoretic formulations[3, 6, 7] of them. To understand the importance of the absence of gap, one can picture the isotropic to the layered smectic phase transition in lyotropic liquid crystals[8], which is always separated by a density jump, just like the liquid-solid transition, and this is in sharp contrast with conditions for the existence of a Lifshitz point. The fact that one can go continuously (without gap) from the disorder phase to the layered phase, implies some constraints on the type of the underlying molecular organisation, in particular the existence of stable cluster phases[9], which are intermediates states between fully disordered and fully layered, such as raft-phases, for example. The lambda-line[3] often separates the disorder phase from this cluster phase. This way, when going from the disordered phase into the layer phase, the small wave vector part of the scattered intensity raises first, due to the fluctuation of forming the clusters, and then, instead of diverging as in the case of an orientationally ordering, the small wave vector part of the scattered intensity starts raising a non-zero pre-peak, witnessing the layer ordering. In short, the Lifshitz point implicitly requires some form of microscopic clustering to exist. However, the question of the requirement of implicit clustering in a Lifshitz point context, remains ill-documented and answered. It seems to be a byproduct of the microscopic interactions, or theoretical formulations. It is this through this question that we would like to re-examine here the concept of Lifshitz point.

In the present paper, we propose to extend this concept of Lifshitz point to that of Lifshitz state. The context for such proposition is the following. In recent works[10], we have emphasized the specific nature and structural properties of the hydrogen bonding induced clustering in the context various binary mixtures, such as aqueous[11, 12, 13], non-aqueous[14, 15] mixtures and ionic liquids[16, 17, 18, 19]. In particular, we have demonstrated, through extensive computer simulations, that the atom-atom pair correlation functions  $g_{ab}(r)$  of the clustering species tend to show long range domain oscillations, which in turn produce a domain pre-peak in the corresponding atom-atom structure factors  $S_{ab}(k)$ , where  $a$  and  $b$  designates specific atoms in molecular species. The presence of such a pre-peak is the principal reason we proposed to designate as “molecular emulsion” all systems exhibiting this structural property[10, 20]. This is in analogy with

micro-emulsion systems, which have a pre-peak in the scattered intensity  $I(k)$ [21]. Unlike micro-emulsions, which exhibit various types of micellar phases, as well as lamellar phases, molecular emulsions are essentially disordered phases. It is then highly unlikely that they would possess a Lifshitz point. However, molecular emulsions and micro-emulsions share a common feature, they are both cluster based phases, with a common hydrogen bonding based hydrophilic/hydrophobic characteristics at molecular level. Therefore, they both share the clustering feature from the disordered side of the phases. As we will show here, the domain pre-peak in  $S_{ab}(k)$  can disappear in certain types of mixtures, such as aqueous and non-aqueous mixtures of diols. Since both types of mixtures contain the same microscopic ordering feature, namely the hydrogen bonding ability, which is at the root of the clustering in molecular and micro-emulsions, one may wonder what would be the origin of this difference. We interpret the absence of pre-peak in  $S_{ab}(k)$  as a signature of the Lifshitz state, where the system is stuck in a disorder without clustering, but which is not a random disorder either, precisely because of the strong hydrogen bonding tendencies. From this point of view, the Lifshitz state represents a novel form of disorder, intermediate between pure random disorder and domain order, in a globally disordered homogeneous liquid. It is one step further to classify different forms of disorder in liquids, which could be of importance in classifying different types of disorders, particularly in the context of soft and bio matter.

This study is motivated by our recent computer simulations of several pure n-diols[22], where we found that, despite an apparent clustering of the hydroxyl groups, which produced chain-like clusters, and subsequent pre-peak in the oxygen-oxygen structure factors, the calculated Xray scattering intensities showed only a weak sign of the usual cluster pre-peak found in scattering experiment of monols[23, 24, 25]. This feature was traced back to the fact that the scattering contributions of the methyl groups of linking alkyl chains produced a variety of intermediate pre-peaks, which tend to overcome the contributions of the hydroxyl groups in the summed total contribution to the Xray scattering[22]. In other words, the alkyl chains linking the hydrogen bonding hydroxyl endgroups, tend to produce an effective disorder, despite a strong hydrogen binding order. How does this disorder-order conflict, typical of these n-diols, get affected under mixing with other hydrogen bonding species? We find that it is this conflict which gives rise to a special form of disorder that we call Lifshitz state. This is the principal reason we have focused our study in linear 1,n-diols. For example, branched diols such as 1,2-propandiol, having one side with strong hydrogen bonding and a non-bonding tail, are likely to produce the same type of domain segregation as 1propanol or tbutanol in water mixtures, hence molecular emulsions with a domain segregation pre-peak.

Ethanediol, also called ethylene glycol in the literature, has been extensively

studied in aqueous mixtures, but less in alcohol and ethanol mixing. The literature for higher diols is equally scarce. Thermodynamic studies[26, 27], diffusion[28], viscosity[29] as well as surface tension[30] measurements, seem to indicate less clustering than in aqueous mono-ol mixtures. The measure of preferential solvation through the Kirkwood-Buff integrals by Y. Marcus[31] shows near ideality of the mixtures of 1,n-diols, as opposed to branched ones. Radiation scattering studies show scattering pre-peak for the aqueous short linear diols mixtures[32, 33], which do not seem to differ conclusively from aqueous-monols mixtures. There are quite a few computer simulation studies[34, 35, 36, 37], which focus mostly on testing various force fields issues and studying short range structural properties. Gubskaya and Kusalik[35] report that, despite hydrogen bonding tendencies, little hydrophobic association is found through the study of direct space correlations. It is noteworthy that one of these studies report calculation of structure factors.

In the remainder of the paper, we first present the molecular models we have used as well as details of the protocol of our computer simulations. This is followed by an analysis of the simulation results. In a final section we analyse the adequacy of the concept of Lifshitz state as well as the heuristic perspectives it opens.

## 2 Theoretical and computational details

### 2.1 Theoretical considerations

One of key property of disordered liquid mixtures is homogeneity. This homogeneity is characterised by the fact that the total density  $\rho = N/V$  and the species densities  $\rho_i = N_i/V$ , are scalar order parameters[38], where  $N = \sum_i N_i$  is the total number of particles, with  $N_i$  particles of each species  $i$ , and  $V$  is the total volume. In inhomogeneous systems, these quantities would be functions which depend on the spatial variables which characterise the inhomogeneity, such as the distance to a wall, for a confined liquid, or the orientation of global order for a liquid crystal. Homogeneous order can nevertheless be further characterised by the measure of the fluctuations[38]  $\delta\rho_i(\mathbf{r}) = \rho_i(\mathbf{r}) - \rho_i$  of the local densities  $\rho_i(\mathbf{r})$  for each of the species at spatial position  $\mathbf{r}$ . Such fluctuations can be measured through the pair correlation functions, defined as

$$g_{ij}(r) = \frac{\langle \rho_i(\mathbf{r}_1)\rho_j(\mathbf{r}_2) \rangle}{\rho_i\rho_j} \quad (1)$$

where the symbol  $\langle \dots \rangle$  designates a statistical ensemble average, and  $r = |\mathbf{r}_1 - \mathbf{r}_2|$ . Similar quantities can be equally defined for various atom-atom correlations  $g_{ab}(r)$  in case of molecular mixtures[38, 39], where  $a$  and  $b$  designate atoms belonging

to molecular species. The atom-atom structure factors are defined as the Fourier transforms

$$S_{ab}(k) = \delta_{ab} + \sqrt{\rho_a \rho_b} \int d\mathbf{r} [g_{ab}(r) - 1] \exp(i\mathbf{k} \cdot \mathbf{r}) \quad (2)$$

Both these functions are a direct measure of the fluctuations occurring in the liquid, and for all distances  $r$  and corresponding wave vectors  $k$ . For example, the peak structure in the  $g_{ab}(r)$  will designate the fluctuations corresponding to the presence of atomic cores of atom  $a$  and  $b$ , with periodicity  $\sigma_{ab} \approx (\sigma_a + \sigma_b)/2$ , where  $\sigma_a$  and  $\sigma_b$  designate the diameter of atoms  $a$  and  $b$ , respectively. The corresponding main peak in the structure factor  $S_{ab}(k)$  will be positioned at  $k \approx 2\pi/\sigma_{ab}$ . Similarly, long range density or concentration fluctuations, such as those occurring at the approach of second order phase transitions, through the development of the Yukawa tail in the correlation function  $g_{ab}(r) \rightarrow \exp(-r/\xi)/r$  when  $r \rightarrow +\infty$ , where  $\xi$  is the correlation length, lead to the appearance of an increase of the peak of  $S_{ab}(k=0)$ , through the well known relation to particle number fluctuation[38]

$$S_{ab}(k=0) = \frac{\langle N_a N_b \rangle - \langle N_a \rangle \langle N_b \rangle}{\sqrt{\langle N_a \rangle \langle N_b \rangle}} \quad (3)$$

The similarity of this expression with Eq.(1) shows the common origin they have with fluctuations in general. These are very general textbook considerations.

However, clustering and micro-heterogeneity equally lead to specific long range domain-domain correlations, with a corresponding pre-peak in the structure factor  $S_{ab}(r)$ . We have shown several example of such pre-peak, in particular for the correlations  $S_{OO}(k)$  between the oxygen atoms of the hydroxyl groups in hydrogen bonding mixtures. Since these pre-peak occur at a non-zero wave vector  $k_p \approx 2\pi/d$ , where  $d$  is the average domain size, one must not confuse micro-heterogeneity with the thermodynamic definition of concentration fluctuations as expressed through Eq.(3). These latter can be measured through the Kirkwood-Buff integrals (KBI) [40], which are defined as the integral of the correlation functions  $g_{ab}(r)$  as:

$$G_{ab} = 4\pi \int_0^\infty dr r^2 [g_{ab}(r) - 1] \quad (4)$$

and are related to the  $S_{ab}(k=0)$  through the expression

$$G_{ab} = \frac{S_{ab}(k=0) - \delta_{ab}}{\sqrt{\rho_a \rho_b}} \quad (5)$$

which shows clearly that they are a measure of thermodynamic fluctuations. Therefore, micro-heterogeneity, which concerns the pre-peak of  $S_{ab}(k \neq 0)$  should not be confused with the Kirkwood-Buff integrals, which concern  $S_{ab}(k=0)$ . These

latter quantities have been subject to various confusions in the past literature, in particular with the concept of local solvation [?, ?, ?], which tend to indicate that connection between fluctuations and local structure of complex liquids deserves further developments and clarifications[10].

The KBI can be related to thermodynamic quantities[40], such as the partial molar volumes, the volume and the mole fraction derivatives of the chemical potentials through the expressions[41]

$$G_{ij} = (1 - \delta_{ij}) \left[ \kappa_T^* - \frac{\bar{V}_i \bar{V}_j}{VD} \right] + \delta_{ij} \left[ G_{12} + \frac{1}{x_i} \left( \frac{\bar{V}_j}{D} - V \right) \right] \quad (6)$$

where  $\kappa_T^*$  is the isothermal compressibility (in reduced units, which we often neglect since it is small for dense incompressible liquids),  $\bar{V}_m$  is the partial molar volume for species  $m$ ,  $V$  is the total volume (which we approximate through the linear relation  $V = (1 - x)\bar{V}_1 + x\bar{V}_2$ ) and  $D = D(x)$  is given by[41]:

$$D(x) = -(1 - x) \frac{\partial \beta \mu_1}{\partial x} = x \frac{\partial \beta \mu_2}{\partial x} \quad (7)$$

where  $\mu_2$  is the chemical potential of the diol species 2 in our convention ( $\beta = 1/k_B T$  is the Boltzmann factor, with  $k_B$  the Boltzmann constant and  $T$  the temperature).

This preliminary introduction to the rich nature of fluctuations in mixture is needed to better appreciate the absence of it in the particular cases we propose to examine below.

## 2.2 Models and simulation

We study aqueous mixtures of ethanediol and 1,3-propanediol, as well as ethanol mixture of the two same diols. For each type of binary mixture we have studied 3 diol mole fractions of  $x = 0.2, 0.5$  and  $0.8$ , in addition to the respective pure liquids,  $x = 0$  and  $x = 1$ , which we have studied independently previously[22, 42, 43]. We have used the SPC/E model for water[44], the TraPPE model for ethanol[45] and the two diols[46]. System sizes of  $N=2048$  and  $N=16000$  particles have been studied, the larger system in order to clarify the nature of the long range correlations in these systems, which is an important problem for these mixtures, as we shall discuss below. The initial configurations were generated by random molecular positioning, with the program PACKMOL[47]. The GROMACS code[48] was used for the molecular dynamics simulations, as in many of our previous works. Initial configurations were first energy minimized, then simulated in the constant NVT ensemble for few hundred picoseconds. Then, 1-2 ns constant NPT runs were performed to stabilise the system at ambient conditions, with  $T=300\text{K}$ . Temperature was maintained constant through the Nosé-Hoover thermostat[49, 50],

and pressure was maintained at 1 atm with the Parrinello-Rahman barostat[51, 52], with time constant of 1 ps. Production runs were performed in the ambient conditions, for 2-10 ns, depending on systems and convergence of the long range tail of the atom-atom correlations functions, which is a particularly difficult issue in these mixtures.

By the definition of Eq(1), the atom-atom correlation functions should have an horizontal asymptote at very large separation. This asymptote is 1 in the thermodynamic limit  $N \rightarrow \infty$  and  $V \rightarrow \infty$ . In finite size systems, however, this is never achieved, as demonstrated by Lebowitz and Percus[53] and the asymptote is  $1 - \varepsilon/N$ , where  $\varepsilon$  is related to the concentration fluctuations[53, 54]. In many cases, particularly for pure systems, this shift in the ideal asymptote is visible, and can be corrected by empirical methods of shifting, which we have demonstrated in previous works[55]. When domain-domain long range oscillations exist, they tend to mask the asymptote and this is a problem to determine the proper value of the shift. In the case of the diols mixtures we have studied, we found that the asymptotes were always void of domain oscillations, but it was difficult to obtain flat asymptotes. This will be illustrated below. In some case, we found necessary to perform large scale and lengthy simulations to get better behaviour of the asymptote. We believe that this problem is inherent to this type of system, with sluggish equilibration and statistics, due to the fact that the hydroxyl groups of the diols are tied by the alkyl chains. Then the hydrogen bonds between the various hydroxyl groups tend to slow down the overall refresh of the statistics. The only cure seems to perform very long runs.

### 3 Results

In our previous computer simulation study of pure n-diols[22], we found that these types of alcohols, with the two hydroxyl groups constrained by a linking alkyl chain, produced a chain-like clustering similar to that found in mono-ols[56, 57, 58], as well as corresponding pre-peak in the oxygen-oxygen structure factors, again similarly to mono-ols[58]. However, in contrast to the apparent pre-peak found in the experimental Xray scattering of mono-ols[23, 24, 25], we observed only a weak shoulder in the calculated Xray scattering intensities[22]. The question which underlies this study is about how this dual property of these diols, namely the hydrogen bond induced microscopic order and the contrasting apparent disorder in the experimental observables, will affect the usual domain segregation patterns found in aqueous and alcohol mixtures of these n-diols?

### 3.1 Snapshots

We first illustrate through snapshots the striking and unexpected homogeneous randomness of these mixtures, for the case of equimolar mixtures and for the system size  $N=16000$  particles. Fig.1 shows typical snapshots, with two colouring conventions. The upper row is for the water-ethanediol equimolar mixture. In Fig.1a we show each of the species with different colors, which help visualise species segregation. The snapshot shows that the species are somewhat segregated into small chain-like clusters, but that the overall distribution is quite homogeneous. This is particularly striking in the case of water, because in all our previous works, we always found that water tend to self segregate into very large pockets[10], particularly well visible for equimolar mixtures. We attribute the present apparent homogeneity to the fact that each diol has two hydroxyl groups to bind with, which increases the water-diol hydrogen binding, but also reduces the water self-binding. Fig.1b shows the same snapshot through the same angle, but the hydroxyl groups and the alkyl groups are colored differently. This permits to observe that the mixing is dominated by the aggregation of the hydroxyl groups, but that this aggregation is quite random, similar to the randomisation of the alkyl groups. The lower row show similar snapshots for ethanol-ethanediol mixtures. The visual observation of the small clustering and the apparent homogeneity of these systems explains many of the structural features which we discuss below for each of the mixtures studied.

### 3.2 Clusters

In order to confirm the visually homogeneous appearance seen in the snapshots of the previous section, we have calculated the cluster distribution of various atoms, and in particular the oxygens of the hydroxyl groups which are responsible for the hydrogen bonding, which underlies the formation of clusters and segregated domains. From our previous cluster calculations, if the hydroxyl groups tend to preferentially self-bind in particular structures, such as chains or loops, then one should see a peak in the corresponding cluster distribution. For example, pure alcohols tend to form chain-like aggregates, which produce a peak in the cluster distribution[43, 58]. Pure diols also form chain-like clusters, which produce a cluster peak in the distribution[22], but it is weaker than in linear alkanols, which means there are less such clusters in diols than in mono-ols. Similarly, ethanol in benzene tends to form clustered domains, at the center of which the hydroxy groups form various loop-like clusters, which are equally detected through a peak in the cluster distribution[59]. This peak is absent in aqueous-alcohol mixtures, despite their strong micro-heterogeneity[10], because the alcohol molecules bind preferentially with water and consequently form fuzzy clusters[59] of all sizes.

Fig.2a-d show the cluster distribution of the oxygen atoms inside each of the 4 mixtures (diol oxygen atoms in the main panel, and solvent oxygen atoms in the insets), and for the 3 different concentrations. It is seen that all these curves show a decaying distribution, with no specific peaks. This does not mean that there are no chain clusters, but such clusters are not dominating the distribution, as in the case of pure diols and mono-ols. These distributions are similar to that of the carbon atoms of the alkyl groups (not shown), the latter which are supposed to obey random distribution since they are not hydrogen bound. This is a direct confirmation that there are little specific hydrogen bond based clusters. But, it does not exclude micro-heterogeneity, as seen in the case of aqueous-alcohol mixtures[59]. The clustering distribution of the oxygens of ethanediol presents a peculiarity which is not seen in longer diols: the odd atom clusters are much less probable than the even ones. This asymmetry is illustrated by showing the odd atom distribution as thinner dotted lines in Figs.2a and Figs.2b. The short alkyl chain constraint imposes a form of hydrogen bonding since the end chain hydroxyl groups are strongly correlated to the nearing methylene groups. However, as we will see below, this constraint is not apparent in the correlation functions and the structure factors, because of the overall disorder of the hydrogen bonding.

### 3.3 Structure functions

In order to detect micro-heterogeneity, we need to look at the long range correlations between the oxygen atoms, and the associated pre-peak in the corresponding structure factors. In a previous paper[59], we distinguished between the short range depletion in the correlation functions, past the first peak of  $g_{ab}(r)$ , which is a signature of linear chain-like clusters involving atoms  $a$  and  $b$  (usually the oxygen atoms in case of hydrogen bonded systems), and the long range domain oscillations, which are a signature of micro-segregation. The first type of correlations is seen for example in neat alcohol or ionic liquids[10, 60], while the second type is seen in micro-heterogeneous mixtures. These two types of clustering produce two distinct pre-peaks. The chain correlation produces a pre-peak usually in the  $k_P \approx 1\text{\AA}^{-1}$  range, while the domain pre-peak is for  $0 < k_D < 1\text{\AA}^{-1}$  range.

Following this remark and what we have obtained so far for the present systems, we expect chain pre-peaks, but no domain pre-peaks.

#### 3.3.1 Aqueous-Ethanediol mixtures

Fig.3a shows specific oxygen-oxygen correlations, and Fig.3b the corresponding structure factors, for the three diol mole fractions we have studied. Pure liquid correlations are equally shown in black lines. It is generally seen that water oxygen-oxygen correlations are stronger than the corresponding diol correlations, and they

behave in opposite manner: for water, they become stronger with *decreasing* water content, while for the diol they decrease with decreasing diol content. This is a generic behaviour of aqueous mixtures, which we have reported previously in various contexts. It is a direct consequence of the higher charges on water model oxygen and hydrogen sites than on the hydroxyl groups of the diol solute model, which mimic hydrogen bonding through classical Coulomb interaction. At lesser water content, water always prefers to self-bind to itself rather than to the dominating amount of surrounding solutes. It is seen that the long range correlations are rapidly screened, hence there are no long range domain order.

Since the snapshots reveal small chain-like clusters, we expect to see these through a pre-peak in the oxygen-oxygen structure factors. Fig.3b shows that such pre-peak -or plateau, exists for the diol and cross structure factors (left and middle panels), but the water structure factor (right panel) shows rather an enhanced  $k=0$  peak. The absence of a clear cluster pre-peak for water is very intriguing. This is actually a capital point for the concept of Lifshitz state. It seems that the water correlations “hesitate” between a cluster pre-peak and a concentration fluctuation  $k = 0$  peak. We come back to this point later.

The increase of the  $k = 0$  peak is coupled to the behaviour at larger  $k$ -values. For example, it is interesting to see that both the  $k = 0$  and the cluster pre-peak of the oxygen correlations of the diol (left panel) increase when the diol content decrease (for  $x = 0.2$ ), which is also accompanied by a decrease of the main peak at  $k_M \approx 2.7 \text{ \AA}^{-1}$ . A similar effect is also observed for the water oxygen correlations.

### 3.3.2 Aqueous-Propanediol mixtures

Fig.4a and Fig.4b show the same quantities as on Fig.3a-b, but for aqueous-propanediol mixtures. The overall behaviour of the correlations are very similar to that observed in Fig.3a-b. We see that the real space correlations show a higher first peak for aqueous-propanediol than for aqueous-ethanediol. Since the charges on the oxygen atoms are the same between the 2 models, the increase in the oxygen correlations comes from the presence of additional methylene group for the propanediol. The cluster pre-peak in the structure factor for the oxygen atoms of the propanediol are better defined than for ethanediol, which was also the case for the pure diol[22]. This is probably another indirect effect of the additional methylene groups, which enforce oxygen atom correlations through hydrophobic effects. We again observe that the water oxygen atom structure factors show no sign of any clear pre-peak, and rather a  $k = 0$  increase. For the case of  $x = 0.2$ , we have show a comparison with the  $N=16000$  particles (yellow dashed curve), where the domain pre-peak type feature vanishes for the larger system, indicating that such feature is a numerical artifact.

### 3.3.3 Ethanol-Ethanediol mixtures

The oxygen-oxygen correlations and structure factors of the ethanol-ethanediol mixtures are shown in Figs.5a-b. Looking at the  $g_{ab}(r)$ , we see that the diol behaves like water: the main peak increases with decreasing concentrations of the diol, while it is the opposite for ethanol. This is very different than what we have observed so far when mixing alcohols with non-associating solutes: the alcohol always behaved like water, since the alcohol molecules would always prefer to hydrogen bond with each other, in the middle of non-bonding solutes. The fact that now the diol behaves like water in presence of ethanol is directly imputable to the alkyl constraint linking the two hydroxyl groups. In fact, this is consistent with water, since one could view water as a “zero-level diol” (with zero alkyl chain).

Since both pure liquids have a pre-peak[22], we expect to see in Fig.5b pre-peaks corresponding to  $k_p \approx 1\text{\AA}^{-1}$ , which is indeed the case. But we also see smaller pre-peaks for smaller k-values. These cannot be compared to the domain pre-peaks, which are always very high, with a magnitude ranging from 5 to 50[10], while those we see here are about 1. Indeed, there is no corresponding medium-to-long-range correlations in Fig.5a. We believe that these small pre-peaks are artifacts of the Fourier transforms, coming from small irregularities in the flat asymptotic region of the  $g_{ab}(r)$ , which are due to slow statistical convergence of these particular systems, which we have mentioned in Section 2.2.

### 3.3.4 Ethanol-Propanediol mixtures

Fig.6a and Fig.6b show the same quantities as on Fig.5a,b, but for ethanol-propanediol mixtures. We again observe in Fig.6a, that the diol correlations behave like water, but with a smaller magnitude. This is probably due to the fact that the alkyl tail linking the two hydroxyl groups of propanediol is longer. The shorter this link, the closer the behaviour with water is expected.

The structure factors in Fig.6b bear a close resemblance with Fig.5b. We note again a small pre-peak for small k-values, which again comes from statistical problems in the asymptotes of  $g_{ab}(r)$ . In case of  $x=0.2$ , the  $N=16000$  data (yellow dashed curve) shows that this artifact tends to disappear.

### 3.3.5 Methyl group correlations

In Fig.7 we show some correlations of the chosen methyl group correlations, mostly to see the important differences with that of the oxygen atoms, the latter which dominate the correlations in these liquids. For the  $g_{MM}(r)$ , we note the smaller amplitude of the first peaks, as compared with those of hydrogen bonding sites, as well as the weaker variation with mole fraction. For the structure factors,

we essentially note that they look much like those of an ordinary Lennard-Jones mixtures. Both observations point to the fact that the methyl groups look essentially weakly interacting and correlating, although they play an essential role in the randomisation and homogeneity of the hydroxyl groups. They point to the extreme asymmetric role of hydrophilic and hydrophobic moieties, although both are needed to explain the complexity of such mixtures.

### 3.4 Kirkwood-Buff integrals

Kirkwood-Buff integrals are evaluated through Eq.(4), where in practice, the upper bound is limited to half the box size  $l_B = L_{box}/2$ . The validity of this calculation requires therefore that the correlation functions have decayed to 1 before the inter-site distance  $l_B$  is reached. For system exhibiting micro-structure, the long range part of the correlation is affected by it, hence larger boxes are required. This problem adds to the inherent LP correction mentioned in Section 2.2. We refer to the next subsection some of the numerical problems met in calculating the KBI of the present mixtures.

Fig.8 shows the Kirkwood-Buff integrals of each of the 4 mixtures studied in the previous sections. It is seen that the KBI of the aqueous mixtures, shown in Fig.8a and Fig.8c, are near ideal, and in good agreement with the experimental data in the case of the aqueous-ethanediol system[31], the only such data we could find in the literature. We note that Geerke and van Gusteren[37] have previously reported the KBI for aqueous ethylene glycol, in the small concentration regime, and also in good agreement with experimental data of Marcus[31]. The ideal KBI, corresponding to the choice  $D(x) = 1$  in Eq.(6) are plotted as full lines. This ideality is an indication that these mixture have low concentration fluctuations, in addition to being very homogeneous. There is more uncertainty in the data for water than the diol and cross KBI. This is a direct consequence of the sluggishness of the dynamics and the statistics.

In contrast, the ethanol-diol mixtures show somewhat weak non-ideality. The ideal KBI with  $D(x)=1$  are plotted as dashed lines. First of all, because of the proximity of the volumes of ethanol and the diols, all the ideal KBI are grouped quite close to each other, as can be seen from Eq.(6). In the case of ethanol-ethanediol, we have calculated more points, using 4ns and 8ns statistics for the  $N=2048$  systems. In order to match the simulated KBI, we have extracted the  $D(x)$  function from these simulated points, by inverting Eq.(6), following a recipe explained in [59]. The KBI were then re-evaluated by using this  $D(x)$  and are shown in full lines in Fig.8b and Fig.8d. The functions  $D(x)$  are plotted in the respective inset and show small deviations from  $D(x) = 1$ , the latter which is the ideal value corresponding to ideal chemical potentials  $\beta\mu_i = \ln(\rho_i)$ . This is the principal reason why we claim that these KBI are nearly ideal. For example,

in the case of aqueous-alcohol mixtures, which are very micro-heterogeneous, the maximum of the water-water KBI often range under 500-10000[41]. The observed near ideality of the KBI indicates that there are very little concentration fluctuations in these mixtures. This is consistent with the visual homogeneity observed in the snapshots.

The fact that aqueous-diol mixtures are more ideal than ethanol-diol mixtures is very non-intuitive, and turns out to be a key feature of these mixtures. Indeed, one would expect that the presence of methyl groups in ethanol would help randomise these mixture much more than in the case of water. Our explanation for this contradicting finding is that the ideality of these aqueous mixture is only apparent, and that it is the result of a competition between the self-aggregation tendency of water, as observed in all other types of mixtures, and the constraint imposed on the diol hydroxyl groups by the linking alkyl chain. In other words, there is a very strong hidden order in the aqueous diol mixtures, which produces an apparent ideality and decrease of concentration fluctuations, without domain segregation. Conversely, the alkyl groups of the ethanol molecules hinder this hidden ordering, which would be there if only the hydroxyl groups were present. This hindering produces an apparent fluctuation and non-ideality, which is seen in the KBI, but also in the sluggishness of the dynamics of these ethanol-diol mixtures.

### 3.5 Influence of system size and statistics

As mentioned previously, these diol mixtures are very sluggish, hence long statistics are required, despite the fact that these systems look homogeneous. We discuss here the problems of obtaining the KBI from numerical integration of the correlation functions in finite size simulations.

Fig.9 shows the effects of system size and statistics on the correlation functions, as illustrated for the case of oxygen-oxygen correlations of the aqueous-propanediol mixture for 20% diol. 2 system sizes are reported  $N_1 = 2048$  and  $N_2 = 16000$ . For system size  $N_1$ , we have accumulated statistics for 1ns (gold curve) and 8ns (blue curve). For system size  $N_2$ , we have accumulated statistics for 0.5ns (red curve) and 4ns (green curve). The main panel shows that, for atomic distances below 10AA, all statistics are indistinguishable. However, the inset demonstrates the deviations in the long range parts. It is clear that the longer the runs, the better the asymptotes stabilize closer to 1. The data shown is unshifted for the LP correction mentioned in Section 2.2. Panel (b) shows the corresponding “running” Kirkwood-Buff integral, defined as:

$$G_{ab}(r) = 4\pi \int_0^r ds s^2 [g_{ab}(s) - 1] \quad (8)$$

This integral should asymptotically reach the value of the correct KBI (namely about  $G_{O_w O_w} \approx 42 \text{cm}^3/\text{mol}$ , as reported in Fig.8a). This figure shows the dramatic differences one expects from poor statistics. All the tails have been corrected for the proper asymptote shift to bring them to the expected value 1 (which is why they are almost horizontal). The deviation is not apparent until  $r \approx 10 \text{AA}$  for most of the data. It is clear that the small  $N_1$  system is nearly appropriate, provided long statistics (8ns) are performed. Even then, one sees that the asymptote is not quite flat. Short runs for the larger system size produce artificial oscillations (red curves), which disappear with 4ns statistics (green curve), and converge to an acceptable asymptote, which permits to extract the KBI value reported in Fig.8a. **The dashed green line drawn through the tail of the RKBI indicates our estimate of the KBI, together with the error bar, which is about  $10 \text{cm}^3/\text{mol}$ .** The resulting KBI value is above the experimental value, but this could well be a model problem, which is an expected drawback. But the correct trend of the KBI, namely the quasi ideality of this system, as seen in Fig.8a, has been properly reproduced by the models and simulations. **Longer statistics lead to somewhat reduce the tail oscillations in the KBI, but only systems larger than  $N_2$  would lead to more precise values.** However, we believe that the present estimate, as shown in Fig.8, demonstrate the correct overall behaviour. Other methodologies than Eq(8), such as that proposed in Ref.[61], which takes into account the scaling of fluctuations with system size for the case of moderate heterogeneity, could be adapted to Lifshitz phase disorder, precisely because of the reduced heterogeneity.

### 3.6 Thermodynamics

In addition to structural properties, it is useful to test the validity of the force field models by comparing with thermodynamic data such as the total density and the enthalpy. In Fig.10 we show the general trends for the molar volumes and enthalpies of the simulated systems, and compare with experiments in the case of the volumes, which are found to be in excellent agreement in all cases. As for the enthalpies, we note that both water and ethanol mixtures of a given diol are quite close to each other. We also note that the enthalpies of pure diols are more negative than that of these two solvents, because of the presence of the 2 hydroxyl groups, which clearly dominate the energetical parts. One can argue that the sampling problems we find in these mixtures are equally due to the fact that the large negative energies contribute to lock local equilibrium and hinder the statistical refreshment of the configurations, enforcing lengthy statistics.

## 4 Discussion and Conclusion

The principal feature which emerges from our study is the essential role played by the constraint of the alkyl chain between the two hydroxyl of these 1,n-diols. These double hydroxyl groups help randomize the water and alcohol solvents, but the presence of the alkyl constraint produces a hidden order, which is not detected by the cluster and correlation function analysis. We infer this hidden order from the curious absence of self-segregation of water, as seen in all other types of solutes, as well as the apparent ideality of the KBI of these aqueous mixtures. When comparing these features with the findings of the ethanol-diol mixtures, we see that the alkyl tails of the alcohol bring a disturbance to this hidden order, responsible for the weak non-ideality seen in the KBI. In all of the previously investigated mixtures with hydrogen bonding species, the predominant feature was the appearance of segregated domain pre-peak in the atom-atom structure factors. The present mixtures, despite hydrogen bonding tendencies, do not possess this pre-peak. The presence of hydrogen bonding interactions imposes a specific binding, but this bonding does not produce sufficient local order to induce a pre-peak. Yet, this is not random disorder. In other words, the absence of pre-peak does not necessarily imply random disorder.

This hidden order is the reason why we propose to consider these mixtures as a new type of disorder. From the absence of domain-domain correlations, and the appealing similarity with the disappearance of the pre-peak in the micro-emulsion when approaching the Lifshitz point from the side of the layer ordered phase, we propose to call this new type of disorder the Lifshitz state. As hinted in the Introduction, through this naming, we propose a unification scheme for molecular and micro-emulsions. Although micro-emulsions are micro-heterogeneous at a larger spatial degree, with larger water and oil domains, they share the same microscopic hydrogen bonding interaction and hydrophobic/hydrophilic competition patterns. Similarly, even though molecular emulsions cannot manifest the same macroscopic ordered phase transitions as micro-emulsions, they both certainly share similarities in the pre-transitional fluctuations, because of the same microscopic origins, such as the hydrophobic interactions, for example. The present study hints to the richness of the underlying disorder, from the molecular emulsion side.

Whereas the Lifshitz point in micro-emulsions can be reached by moving through the phase diagram, the present Lifshitz state is a permanent physical state. It can be “reached” by changing the nature of the solute particle, by breaking the alkyl chain or branching the diol, in which case the resulting mixture would show the same molecular emulsion pre-peak as water-methanol or water-ethanol. This is a fictitious operation for the present case, but in case of chemically induced bonding, this passage from a molecular emulsion to a Lifshitz state could well

occur. Such type of situation is likely to occur in bio material liquids, which have a rich internal kinetics. The concept of Lifshitz state or disorder could be useful in such context.

The type of special disorder found in these mixtures has suggestive analogies to that separating water from a simple disordered Lennard-Jones fluid[42]. Both liquids are disordered, but the disorder in water has a richness, which reveals itself under mixing with other liquids. Micro-heterogeneity is one such property, which is not found in mixtures of simple liquids, such as mixture of alkanes, for example[54]. Through this paper, we have introduced a new type of structured disorder, which is different from micro-heterogeneity and clustering, but which is not a strict random disorder either, and which we call the Lifshitz disorder.

There is a rich variety in the ordered phases, whether these are crystalline or liquid crystalline phases. The richness of these orders are witnessed through the macroscopic order parameter, which is often the 1-body density  $\rho^{(1)}(1)$ , where the argument indicates that this is function of spatial and/or orientational degree of freedom of each molecule. This 1-body function can be formally related to a spontaneous macroscopic field through modern density functional theory[?], and this field orders the whole system. In contrast, the order parameter of a disordered liquid is the scalar density  $\rho$ , which is therefore unable to express the underlying richness of the disorder. Indeed, the order of the macroscopically disordered phase is local. But the local density  $\rho(1)$  is a random variable, whose statistical average wipes out all the interesting richness to give a scalar:  $\langle \rho(1) \rangle = \rho$ . However, pair correlation functions  $\rho^{(2)}(1,2) = \langle \rho(1)\rho(2) \rangle$ , which are in fact a measure of fluctuations, allow to measure this disorder at small molecular separation, or in the small wave vector limit, as we have seen through this study. The measure of the richness of disorder is important in the case of soft matter or bio-matter, which are essentially disordered systems, but with a rich microscopic order. We hope that the concept of Lifshitz state is a first step into classifying the forms of disorders, and that it will be helpful in future studies.

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## Figure Captions

- Fig.1 Snapshots of the equimolar water-ethanediol (upper row) and ethanol-ethanediol (lower row), shown with 2 different colouring conventions. **(a)** water molecules are shown in yellow and diol molecules in cyan; **(b)** oxygen atoms are in red, hydrogen in white and methylene groups in blue; **(c)** molecules are shown in gray and diol molecules in green; **(d)** same colouring convention as in **(b)**.
- Fig.2 Clusters distributions of the diol oxygen sites (main panels) and solvent oxygen sites (insets), versus the cluster size, for **(a)** the water-ethanediol mixtures, **(b)** the ethanol-ethanediol mixtures, **(c)** the water-propanediol mixtures and **(d)** the ethanol-propanediol mixtures. The diol concentrations are shown in red for  $x = 0.2$ , green for  $x = 0.5$  and blue for  $x = 0.8$ . The thin dotted line in **(a)** and **(b)** are explained in the text.
- Fig.3 Oxygen-oxygen correlation functions **(a)** and structure factors **(b)** for the aqueous-ethanediol mixtures. The diol concentrations are shown in blue for  $x = 0.2$ , green for  $x = 0.5$  and red for  $x = 0.8$ . Pure solvent data is shown in black. Left panel for the oxygen atoms of the diol, middle panel for the cross correlations and right panel for the oxygens atom of water.
- Fig.4 Oxygen-oxygen correlation functions **(a)** and structure factors **(b)** for the aqueous-propanediol mixtures, with same colouring conventions as in Fig.3. The dashed yellow curve is for  $N=16000$  particles in case of  $x = 0.2$ .
- Fig.5 Oxygen-oxygen correlation functions **(a)** and structure factors **(b)** for the ethanol-ethanediol mixtures, with same colouring conventions as in Fig.3.
- Fig.6 Oxygen-oxygen correlation functions **(a)** and structure factors **(b)** for the ethanol-propanediol mixtures, with same colouring conventions as in Fig.3. The dashed yellow curve is for  $N=16000$  particles in case of  $x = 0.2$ .
- Fig.7 Central methylene group correlations( **(a)** and **(c)** ) and structure factors( **(b)** and **(d)** ) for the aqueous-propanediol (upper row **(a)** and **(b)**) and ethanol-propanediol (lower row **(c)** and **(d)**). The line colors correspond to different concentrations, with the same convention as

Fig.3. Full lines for the methylene of ethanediol and dashed for that of ethanol.

Fig.8 Kirkwood-Buff integrals, versus the diol mole fractions, of the aqueous-ethanediol mixtures **(a)**, ethanol-ethanediol mixtures **(b)**, aqueous propanediol mixtures **(c)** and ethanol-propanediol mixtures **(d)**. Blue line is for solvent-solvent, green for solvent-diol and magenta for diol-diol KBI. The squares represent the simulation results, the full or dashed lines correspond to different choices of  $D(x)$ , shown in the inset for the ethanol-diol mixtures (see text). The dots in **(a)** are experimental results from Ref[31].

Fig.9 Influence of system size and statistics. (a) Main panel: water oxygen-oxygen correlations  $g_{O_wO_w}(r)$  for 20% propanediol in water ; inset: zoom on the tail (uncorrected). (b) Running KBI. Gold line is for  $N_1 = 2048$  and 1ns, blue for  $N_1$  and 8ns, red line for  $N_2 = 16000$  and 0.5ns, green line for  $N_2$  and 4ns (some lines are dashed for better visibility). **The horizontal dashed line is the final retained value for KBI, with indicative error bar.**

Fig.10 Thermodynamic properties. Volumes in (a) and (b), enthalpies in (c) and (d). Black line for experimental data [62, 63, 64, 65], purple circles for aqueous-ethanediol, gold squares for ethanol-ethanediol, cyan symbols for aqueous-propanediol and red triangles for ethanol-propanediol.

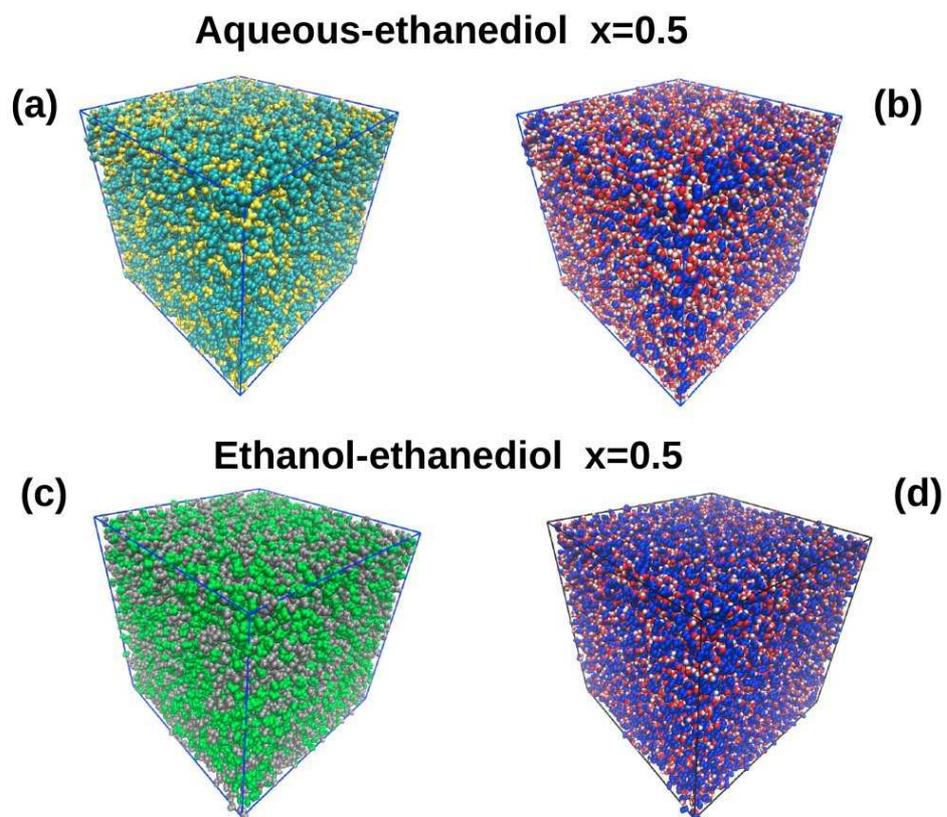


Fig.1 Snapshots of the equimolar water-ethanediol (upper row) and ethanol-ethanediol (lower row), shown with 2 different colouring conventions. **(a)** water molecules are shown in yellow and diol molecules in cyan; **(b)** oxygen atoms are in red, hydrogen in white and methylene groups in blue; **(c)** molecules are shown in gray and diol molecules in green; **(d)** same colouring convention as in **(b)**.

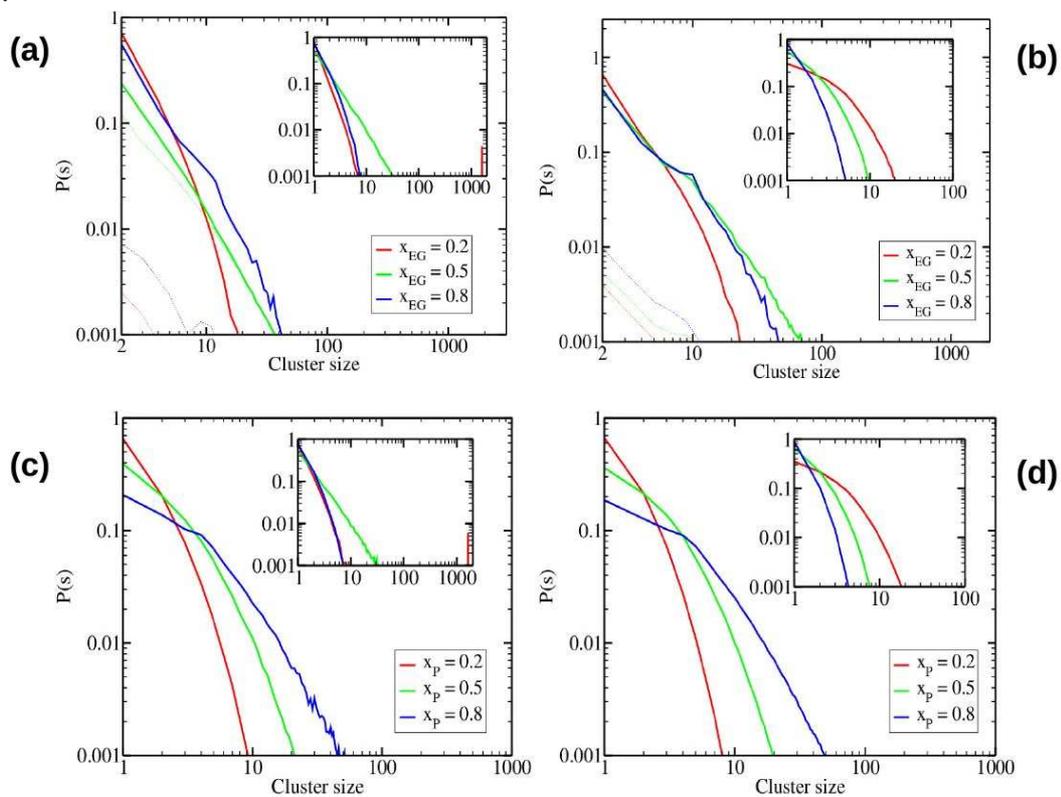


Fig.2 Clusters distributions of the diol oxygen sites (main panels) and solvent oxygen sites (insets), versus the cluster size, for (a) the water-ethanediol mixtures, (b) the ethanol-ethanediol mixtures, (c) the water-propanediol mixtures and (d) the ethanol-propanediol mixtures. The diol concentrations are shown in red for  $x = 0.2$ , green for  $x = 0.5$  and blue for  $x = 0.8$ . The thin dotted line in (a) and (b) are explained in the text.

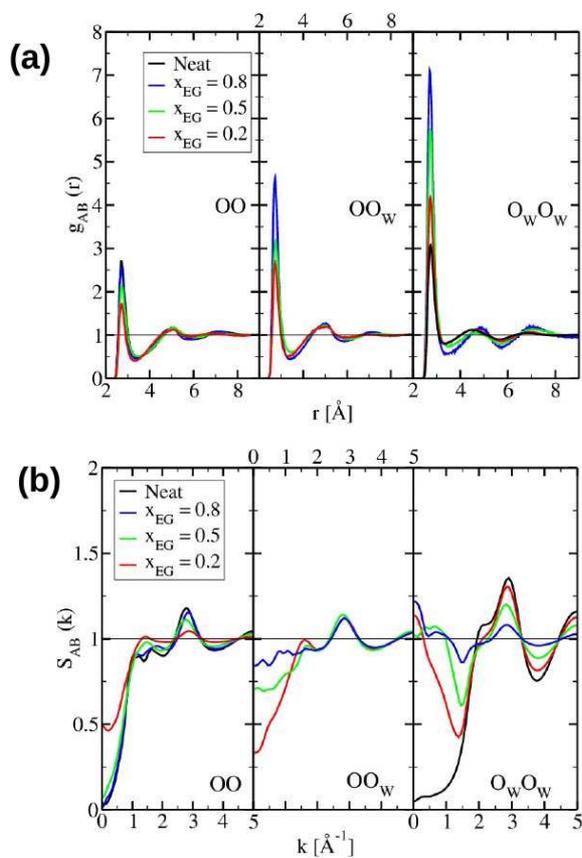


Fig.3 Oxygen-oxygen correlation functions (a) and structure factors (b) for the aqueous-ethanediol mixtures. The diol concentrations are shown in blue for  $x = 0.2$ , green for  $x = 0.5$  and red for  $x = 0.8$ . Pure solvent data is shown in black. Left panel for the oxygen atoms of the diol, middle panel for the cross correlations and right panel for the oxygens atom of water.

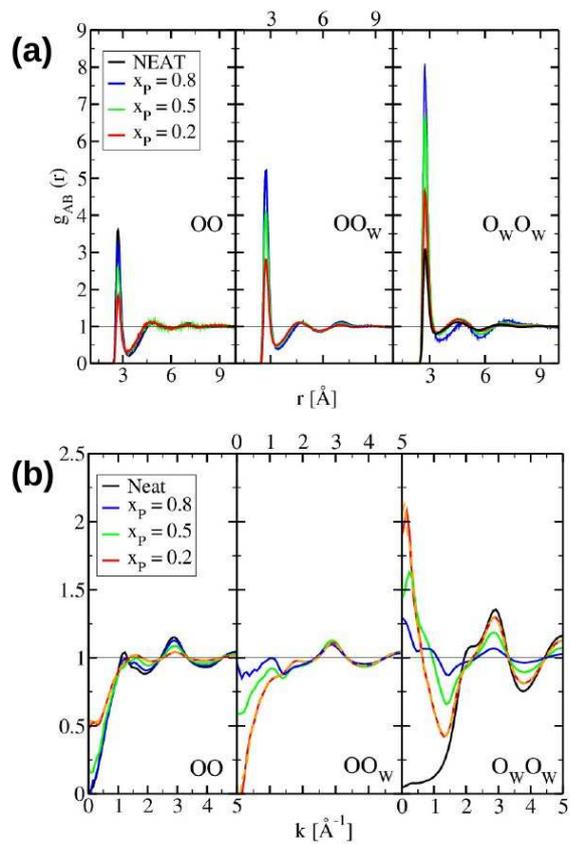


Fig.4 Oxygen-oxygen correlation functions (a) and structure factors (b) for the aqueous-propanediol mixtures, with same colouring conventions as in Fig.3. The dashed yellow curve is for  $N=16000$  particles in case of  $x = 0.2$ .

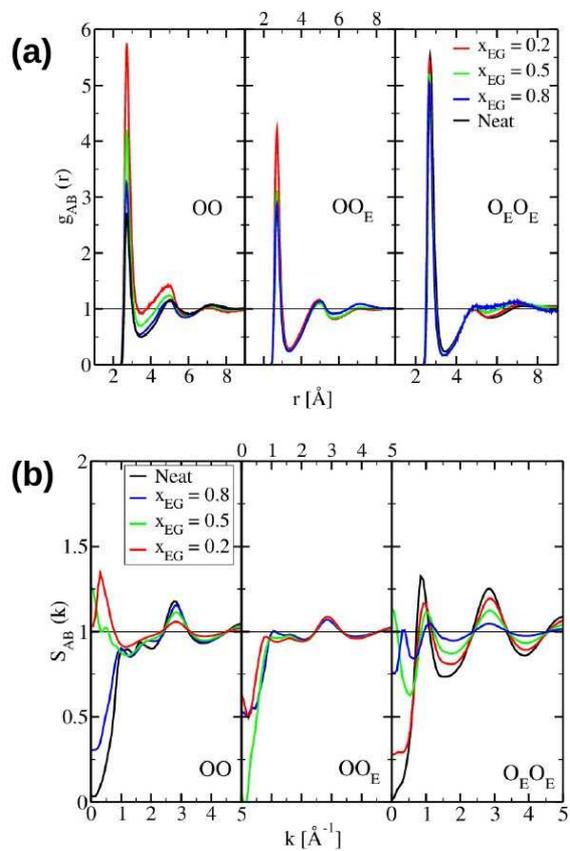


Fig.5 Oxygen-oxygen correlation functions (a) and structure factors (b) for the ethanol-ethanediol mixtures, with same colouring conventions as in Fig.3.

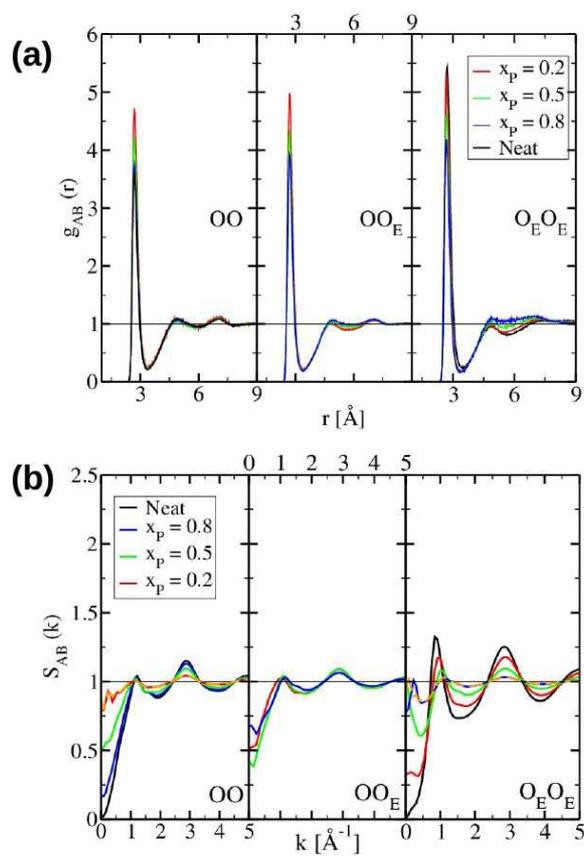


Fig.6 Oxygen-oxygen correlation functions (a) and structure factors (b) for the ethanol-propanediol mixtures, with same colouring conventions as in Fig.3. The dashed yellow curve is for  $N=16000$  particles in case of  $x = 0.2$ .

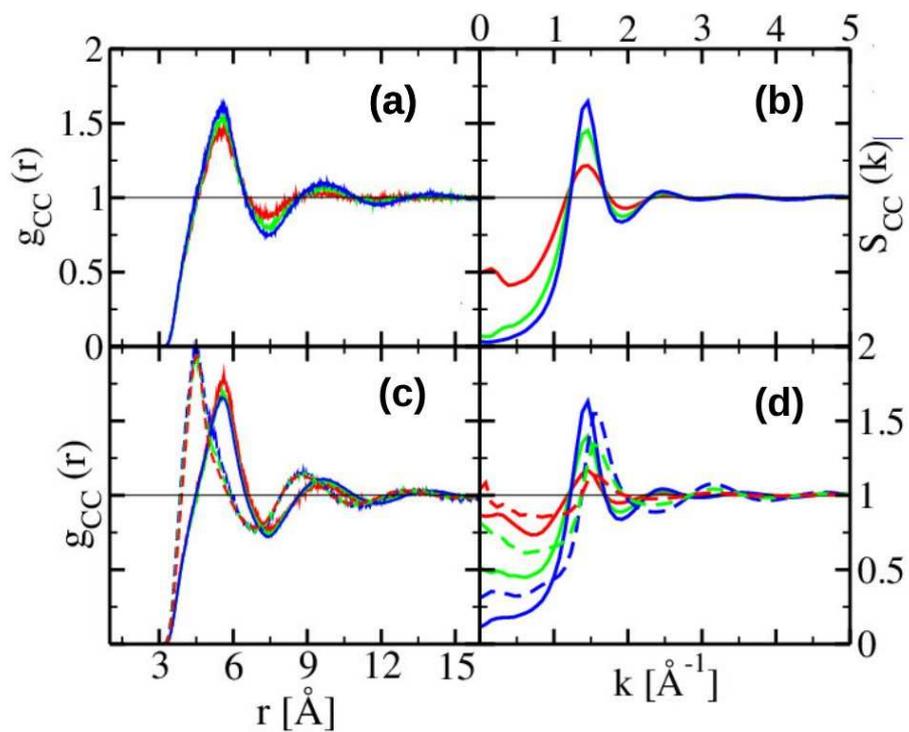


Fig.7 Central methylene group correlations( **(a)** and **(c)** and structure factors( **(b)** and **(d)** ) for the aqueous-propanediol (upper row **(a)** and **(b)**) and ethanol-propanediol (lower row **(c)** and **(d)**). The line colors correspond to different concentrations, with the same convention as Fig.3. Full lines for the methylene of ethanediol and dashed for that of ethanol.

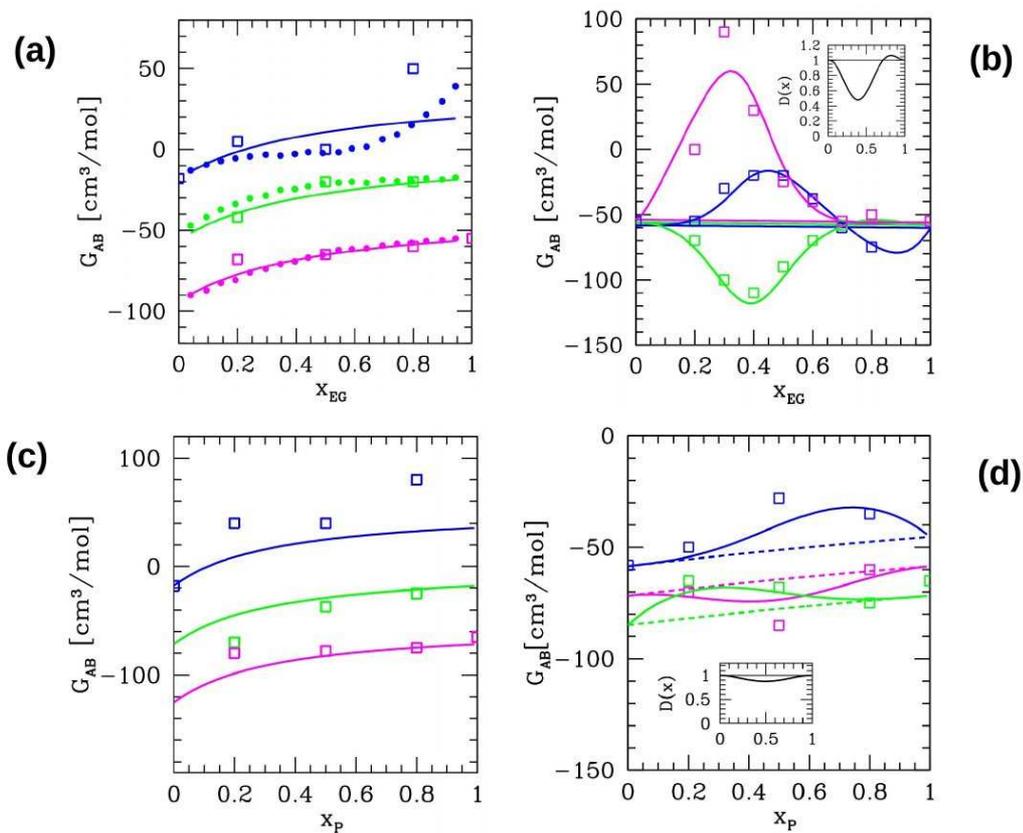


Fig.8 Kirkwood-Buff integrals, versus the diol mole fractions, of the aqueous-ethanediol mixtures (a), ethanol-ethanediol mixtures (b), aqueous propanediol mixtures (c) and ethanol-propanediol mixtures (d). Blue line is for solvent-solvent, green for solvent-diol and magenta for diol-diol KBI. The squares represent the simulation results, the full or dashed lines correspond to different choices of  $D(x)$ , shown in the inset for the ethanol-diol mixtures (see text). The dots in (a) are experimental results from Ref[31].

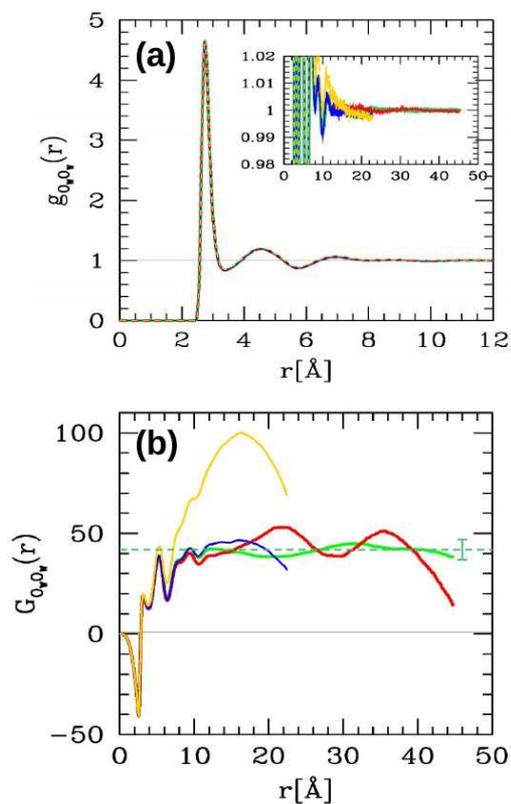


Fig.9 Influence of system size and statistics. (a) Main panel: water oxygen-oxygen correlations  $g_{O_W O_W}(r)$  for 20% propanediol in water ; inset: zoom on the tail (uncorrected). (b) Running KBI. Gold line is for  $N_1 = 2048$  and 1ns, blue for  $N_1$  and 8ns, red line for  $N_2 = 16000$  and 0.5ns, green line for  $N_2$  and 4ns (some lines are dashed for better visibility) .The horizontal dashed line is the final retained value for KBI, with indicative error bar.

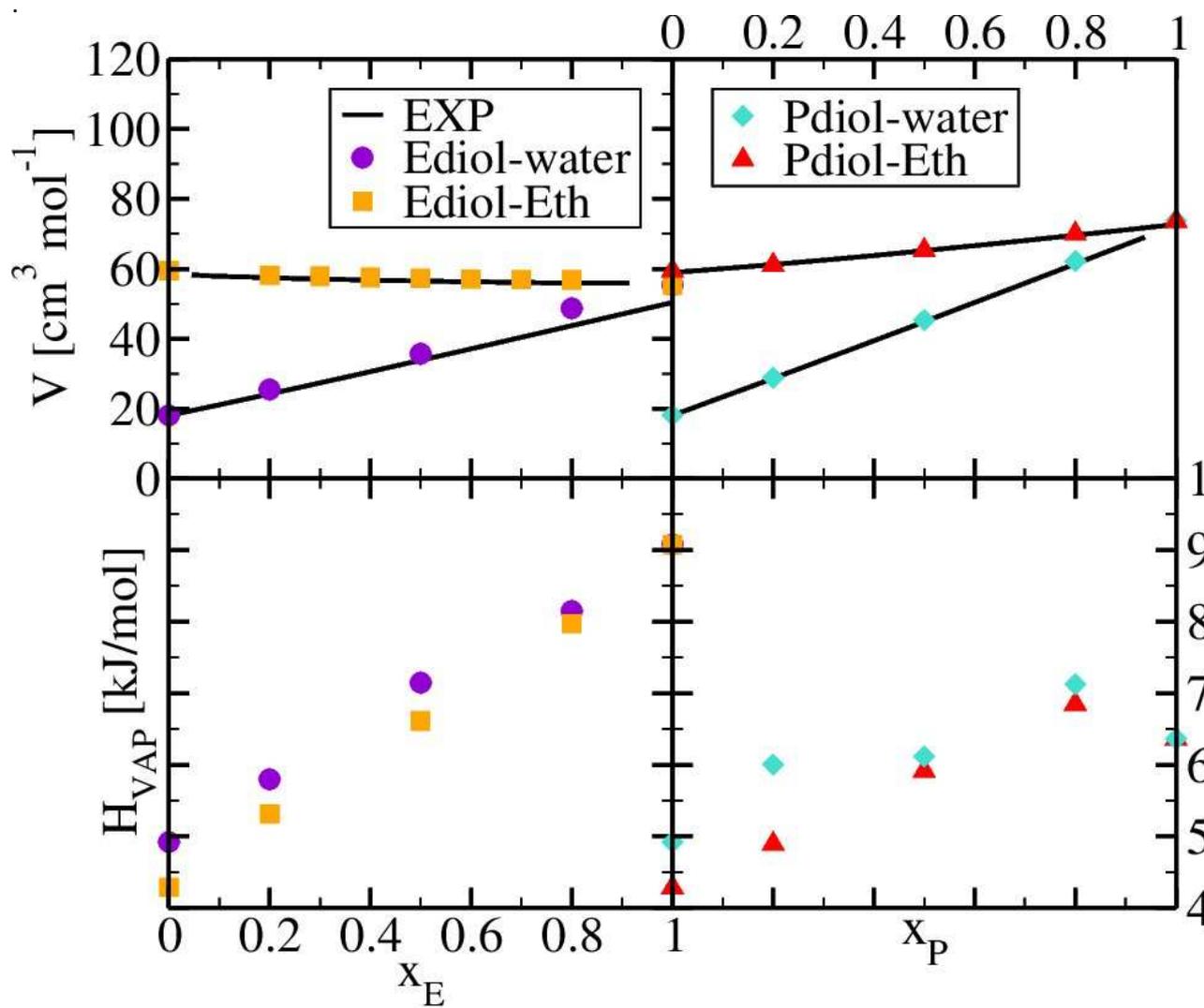


Fig.10 Thermodynamic properties. Volumes in (a) and (b), enthalpies in (c) and (d). Black line for experimental data [62, 63, 64, 65], purple circles for aqueous-ethanediol, gold squares for ethanol-ethanediol, cyan symbols for aqueous-propanediol and red triangles for ethanol-propanediol.