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Fe-N-C electrocatalyst and its electrode: are we talking about the same material?

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ABSTRACT. Evaluation of the electrocatalyst performance data includes an electrode preparation step. Herein, we compare the structural composition of Fe-N-C materials, used to electrocatalyze the oxygen reduction reaction in proton-exchange membrane fuel cells, before and after catalyst

layer preparation. The effects of this step on the electronic structure and local coordination of Fe were investigated by X-ray absorption (XAS) and emission spectroscopies (XES), for Fe-N-C materials prepared via different synthetic routes. This work underlines the importance of determining the Fe-N-C catalyst structure in the prepared electrode for further studies of the structure – activity – stability correlations.

TEXT

The preparation of a homogeneous catalyst layer (CL) for electrochemical testing in either a rotating disc electrode (RDE) or in fuel cell set-up is a critical step. While a considerable amount of work has already been devoted to the preparation and analysis of Pt-based electrodes¹, similar studies on Pt-free materials, such as iron-nitrogen-carbon (Fe-N-C) catalysts, one of the most promising materials for acid and alkaline fuel cells, are less common.²⁻⁴ Fe-N-C electrocatalysts are known to possess a complex composition containing typically several FeN_xC_y moieties (single Fe atoms coordinated to N atoms embedded into a C matrix) of different geometry, where the synthesis protocol favors the presence of certain FeN_xC_y sites and their amount.⁵ Besides, Fe-N-C catalysts might also include Fe-based particles such as metallic Fe, Fe carbide or oxides.⁶ The analysis of the structure of Fe atoms in Fe-N-C materials is often carried out using Mössbauer (MS) and/or X-ray absorption spectroscopies, where other complementary techniques such as electron microscopy, X-ray photoelectron spectroscopy, Raman spectroscopy are used for further characterization of the material both *ex situ*^{7,8} and *in situ/operando*.⁹⁻¹¹ In general, the research related to Fe-N-C catalysts is focused either on the correlations between catalyst activity and its structure defined *ex situ*, or on the analysis of catalyst behavior under reaction conditions, *i.e. in situ/operando*. It is also known that some Fe-N-C materials are sensitive to their storage conditions

(oxygen and water presence)^{2,12,13} as well as to any physical and/or chemical treatment resulting in the electrochemical activity changes.^{14,15} Thus, one may expect some differences in the catalyst structure for a pristine powder and for a CL prepared based on it. In fact, it has been observed in the Mössbauer analysis of Fe-N-C catalyst powders and their CLs that the ratio between different FeN_x sites was changed² and new Fe-based species developed, such as Fe-based oxides.¹⁶

In this work, we applied Fe K β high energy resolution fluorescence detected (HERFD) X-ray absorption near edge structure (XANES) and K β X-ray emission spectroscopy (XES) techniques to analyze the structure of Fe-N-C catalysts in the CLs, *i.e.* Fe-N-C-based electrodes, and as the pristine powders. Analysis of the XANES provides information on the coordination of the Fe atoms, whereas K β XES, comprised of K $\beta_{1,3}$ and K β' components, is sensitive to changes in the Fe total spin.^{17,18} The XAS and XES measurements were performed at beamline ID26 of the European synchrotron (ESRF). The experimental details are provided in the Supporting Information.

The catalysts analyzed in this work have been obtained via different preparations: Fe0.5 is a metal-organic framework (MOF)-derived material prepared via ramp pyrolysis synthesis (labeled as Fe0.5RP in previous publications)¹⁹; C-PPy, is a polypyrrole-derived isotopically enriched ⁵⁷Fe-N-C;²⁰ additionally a non-pyrolysed carbon-supported porphyrin Porph/C_{prec} and a porphyrin-based catalyst Porph/C_{cat} are included, both related to FeTMPPCl (5,10,15,20-tetrakis(4-methoxyphenyl)-21H,23H-porphine iron (III) chloride) supported on carbon.⁸ More details on the preparation procedures can be found in the Supporting Information. On the basis of previous data of low- and room-temperature Mössbauer spectroscopy, one can describe the materials with respect to their Fe-related compositions as follows: Porph/C_{prec} contains only porphyrin-like FeN₄ environments with an axial (chlorine) ligand,⁸ both Fe0.5⁵ and Porph/C_{cat}

might be dominated by FeN_xC_y moieties, whereas C-PPy represents a mixture of molecular FeN_xC_y sites and crystalline or amorphous side phases.^{8,10}

The Fe $\text{K}\beta$ HERFD XANES and XES spectra recorded on the catalysts powders are shown in Figure 1 and Figure S1. XANES signatures of all analyzed catalysts possess characteristic features generally attributed to the FeN_xC_y sites.^{5,11} The energy of the first inflection point of the Fe K edge is increasing in the following order: Porph/ C_{prec} (7122.6 eV) < Fe0.5 (7123.9 eV) < C-PPy (7124.8 eV) < Porph/ C_{cat} (7125.6 eV). Taking into account that the edge position is affected by several factors, *i.e.* Fe coordination number (CN), coordinated ligand nature and Fe valence, discrimination between Fe^{2+} and Fe^{3+} cannot be done only based on the position of Fe K edge.²¹ For comparison, the first inflection point for Fe_2O_3 (with majority of Fe^{3+}) is at 7122.4 eV, lower than the position of the first inflection point for some of the studied catalysts, that are known to be a mixture of Fe^{2+} and Fe^{3+} oxidation states. This can be explained with the different ligand nature (oxygen / nitrogen / carbon / chlorine) and Fe coordination^{22,23} without invoking a change in average oxidation state of iron atoms. The pre-edge and rising edge features at ≈ 7114 eV and 7119 eV, respectively, are sensitive to the Fe local geometry and its electronic structure. The rising edge (≈ 7119 eV) is characteristic of centro-symmetric square-planar ($\text{D}_{4\text{h}}$) Fe^{2+} components^{11,24} and is well pronounced for Fe0.5, C-PPy and Porph/ C_{cat} , however it is absent in the case of Porph/ C_{prec} containing mostly 5-fold coordinated Fe sites. Considering FeN_4 molecular environments present in the catalysts Fe0.5, C-PPy and Porph/ C_{cat} , the pre-edge peak (≈ 7114 eV) is assigned to 4-fold Fe^{2+} or Fe^{3+} coordinated environment of the FeN_xC_y sites.²⁵ Interestingly, the pre-edge position of square-pyramidal ($\text{C}_{4\text{v}}$) sites of Porph/ C_{prec} is at the same position, while its intensity is slightly higher than for Porph/ C_{cat} (Figure S1) due to its

noncentrosymmetric environment.²⁵ No significant difference in the K β spectra of the catalysts powders is observed indicating that the total spin on Fe does not change.

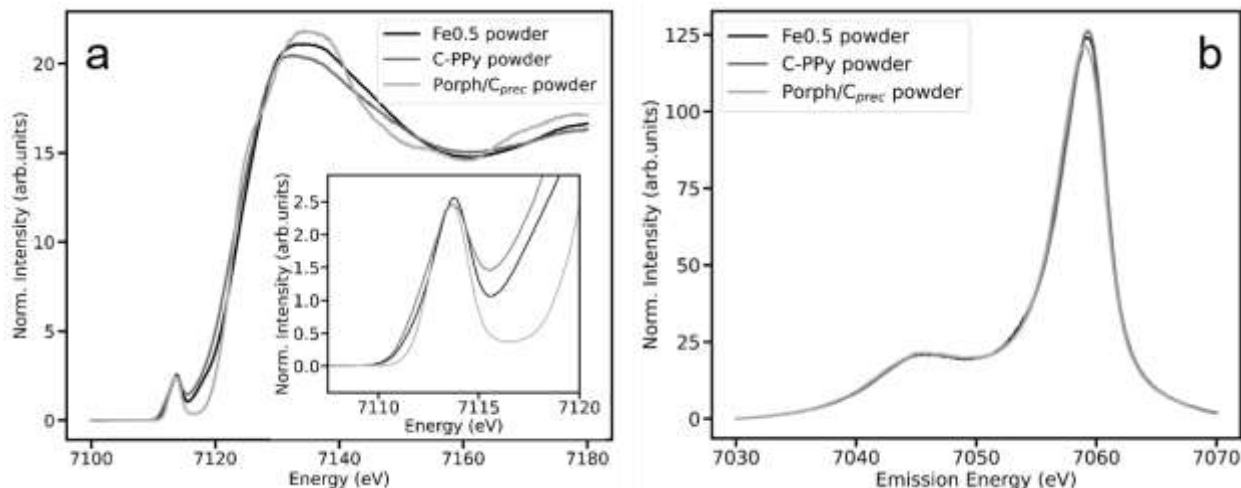


Figure 1. Fe K β HERFD XANES (a) and XES (b) spectra recorded on Fe-N-C catalyst powders. The pre-edge region is shown in the inset of panel a.

Next, we compare the XANES and XES signatures of the Fe-N-C powders and their CLs (Figure 2, Figure S1), where one can clearly see changes in the pre-edge region and in the white line (WL) of the Fe K edge as well as in K β emission line. For all materials, we observe qualitatively the same trends when going from catalyst powder to their CLs: (i) increase in the WL intensity for both peaks at ≈ 7131 eV and ≈ 7139 eV; (ii) shift of the Fe K edge position towards higher energies; (iii) decrease of the pre-edge peaks intensity; (iv) higher low-energy component of K β , K β' , intensity and (v) blue shift of the K β mainline. Interestingly, the magnitude of change differs for the various catalysts. Porph/C_{prec} and Porph/C_{cat} (Figure 2c and Figure S1) seem to be least affected by the CL preparation procedure, whereas Fe0.5 shows the largest changes both in the XANES and XES data.

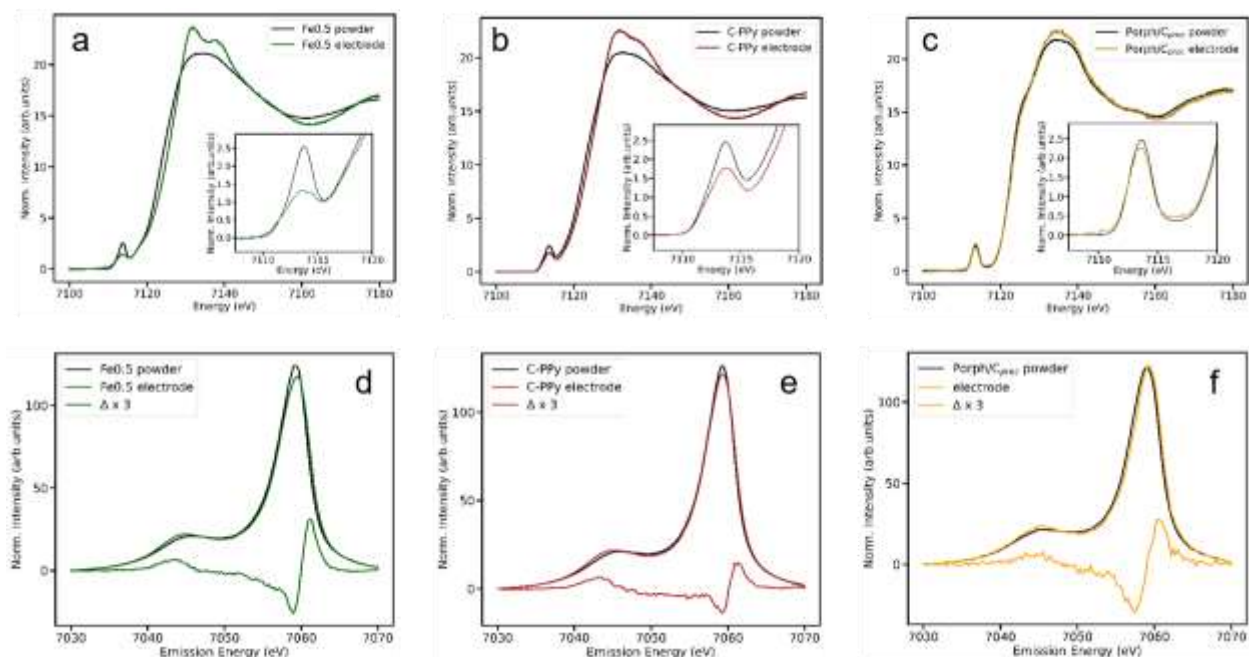


Figure 2. Comparison of Fe K β HERFD XANES (a-c) and XES (d-f) data obtained on the catalysts powders and their electrodes: for Fe0.5 (left), C-PPy (middle) and Porph/C_{prec} (right).

Pre-edge regions are shown in the insets of the corresponding panels. Magnified Fe K β difference signals (Δ) for powder and electrode are shown in the bottom of the panels d-f.

Composition of the catalyst inks used in this work for preparing Fe-N-C-based CLs are listed in Table S2 and included water, isopropanol and Nafion[®] solution (which itself contains water, ethanol and isopropanol). Previously, Artyushkova *et al.* demonstrated a strong interaction of Nafion ionomer with an Fe-N-C catalyst in the CL.³ Also Kramm *et al.* found a significant difference in absorption behavior for the as-prepared CL compared to the Fe-N-C catalyst.² Based on these and other studies^{26–28}, the influence of the Nafion-to-catalyst ratio (NCR) on the resulting Fe-N-C structure in the prepared electrodes was investigated by analyzing their spectroscopic signals. As the spectroscopic changes are qualitatively similar for all the catalysts discussed in this work, we selected MOF-based catalyst Fe0.5 for this NCR study as it shows the most significant changes. Figure 3 presents Fe K edge and Fe K β spectra for Fe0.5 powder and its CLs with the following NCR ratios: 0; 0.25; 0.5; 1 and 2. The increase of Nafion content from 0 to 2 results in

a gradual shift of the Fe K edge, decrease of the pre-edge peak and growth of the white line as well as increase of K β ' peak intensity accompanied with K β mainline shift towards higher emission energies (Table 1). Interestingly, we observe the same changes but less pronounced in the CL prepared from Nafion-free ink (NCR = 0).

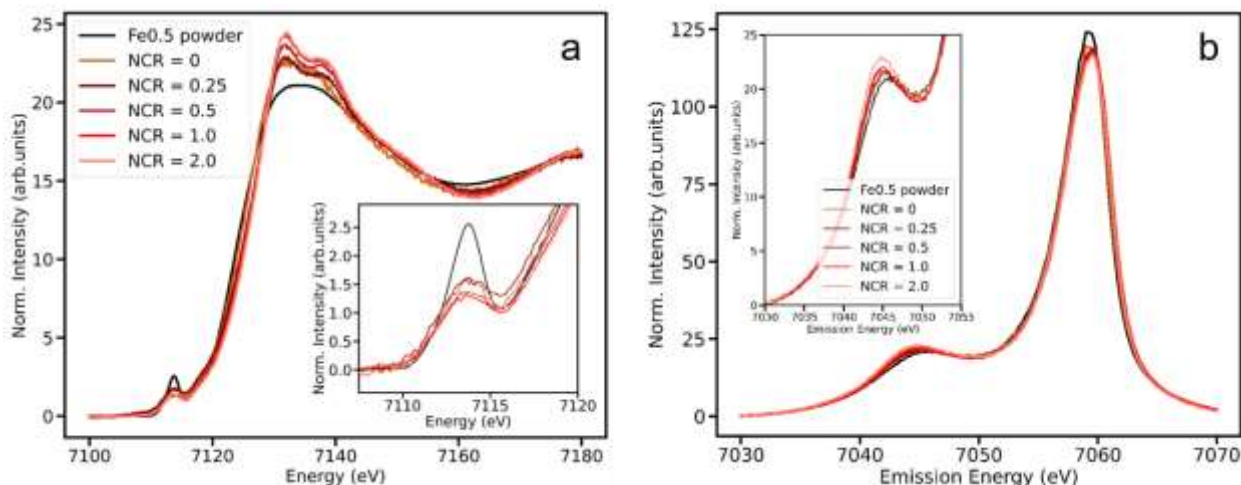


Figure 3. Fe K β HERFD XANES (a) and XES (b) spectra recorded on Fe0.5 catalyst powder and its electrodes with various NCR ratio. The insets show magnified pre-edge (a) and K β ' (b) regions.

Next, we discuss the possible reasons for the observed spectral changes (summarized in Table 1) taking place during or after CL preparation as well as the role of Nafion in the CL. XANES spectra are affected by the local structure (bond distances and angles), ligand nature and oxidation state, where several descriptors define each spectral feature.²¹ The pre-edge region spectral shape depends on the number of electrons in the *d*-shell, where its centroid and peak shape are sensitive to the oxidation state of Fe, while the pre-edge intensity is proportional to the amount of $3d - 4p$ hybridization, *i.e.* it is sensitive to site symmetry. The decrease in the peak intensity at ≈ 7114 eV, observed for all CLs studied here compared to the catalyst powders (Table 1), can be explained by an increase of the Fe coordination number via a square planar (D_{4h}) \rightarrow octahedral (O_h) or pyramidal (C_{4v}) \rightarrow O_h transitions by formation of axial iron-oxygen/hydroxide bonds

during/after CL preparation.²⁵ Based on the XAS and XES data discussed here, one cannot distinguish the exact nature of the bonded ligands, moreover, depending on the local pH, the octahedral sites can be formed via adding --OH...H or H-O-H groups in case of mild acidic solutions. No changes in the pre-edge centroid is observed for all the studied materials suggesting that the Fe oxidation state does not change. It has been shown for Fe-based minerals of the same oxidation state that the edge position of the O_h structures is typically higher than for D_{4h} / C_{4v} .²² The analysis of the 1st inflection point of Fe-N-C-based CLs shows a steady edge shift towards higher energies in comparison to the catalyst powders that can be attributed to changes only in Fe coordination at the iron-nitrogen sites. It should be noted, that not all the sites are expected to transform but mostly the surface- and/or sub-surface-located ones that can be easily accessed by oxygen/water molecules. Additionally, the increase in CN resulting from additional oxygen ligands would influence the white line intensity and lead to sharpening of the XANES features,²¹ as observed for the analyzed materials. Changes of the local geometry will also affect the spin state of Fe atoms. The Fe $\text{K}\beta$ spectrum is governed by $3p - 3d$ exchange interactions and is sensitive to the Fe spin state.¹⁸ The $\text{K}\beta$ differences Δ between powder and CL shown in the bottom of panels d-f of Figure 2 and Figure S4 exhibit shape that is typical for an increase in total spin.¹⁷ A quantitative analysis of the $\text{K}\beta$ spectra can be carried out using the $\text{K}\beta_{1,3}$ first moment (M_1), representing the overall blue shift of the emission line, and by the integrated absolute difference (IAD) method (see the details in the Supporting Information).¹⁷ Independent of the Fe-N-C catalyst, we observe an increase in the IAD value for the CLs in comparison to the powders. The 4-fold/5-fold \rightarrow 6-fold transition proposed earlier based on XANES data explains the increase of the IAD value due to a weaker mixing between ligand and metal orbitals in case of O_h coordination in comparison to D_{4h} and C_{4v} as it was discussed in Ref. ¹⁷. Thus, all XAS and XES changes

discussed here can be explained with a partial formation of octahedral sites during the CL preparation, where the degree of transition depends on the catalyst initial structure.

The analysis of Nafion contribution to the formation of $\text{--O}_2\text{--OH}$ bonds is presented in Figure 4, where four parameters were tracked in the Fe0.5-based CL layers: $\text{K}\beta$ IAD values, 1st inflection point of XANES, pre-edge and white line intensities. Interestingly, the observed changes are not triggered by Nafion, but rather by the presence of either water, oxygen and/or isopropanol in the catalyst ink. However, the introduction of Nafion in the catalyst suspension does enhance them reaching a saturation at ca. NCR of 1, probably due to higher water concentration in the CL resulting from the water vapor adsorption by Nafion ionomer clusters formed in the ink.

The formation of iron oxides during Fe-N-C-based CL preparation has been proposed earlier in literature.¹⁶ For this reason, we compared XANES and XES data, where the valence-to-core (VtC) region was also included, of the analyzed catalysts with the ones' recorded on a commercial bulk $\alpha\text{-Fe}_2\text{O}_3$ powder (CAS number: 1309-37-1) and also on a home-made Fe nano-oxides (6-8 nm according to the TEM images) deposited on N-doped carbon powder ($\text{Fe}_2\text{O}_3/\text{N-C}$)¹⁶. The VtC spectrum arises from transitions from filled valence orbitals that are dominantly ligand in nature where the low-energy VtC feature, $\text{K}\beta''$, is known to be a diagnostic of the ligand bonded to the metal atom.²⁹ For both iron oxides we observed a similar XANES and VtC signatures (Figure S2 and S3) that is consistent with the spectra available in literature.³⁰ Particularly, the pre-edge region (Figure S2b) possesses a well-resolved doublet (≈ 7112.8 eV and ≈ 7114.2 eV) and the low energy VtC component, $\text{K}\beta''$, at ≈ 7091 eV (Figure S3) arises from Fe-O bonds. None of these features is observed in the spectra of the CL and we thus find no evidence for the formation of a significant amount of iron oxides during the CL preparation.

Table 1. Characteristic parameters of the K β XES and XANES spectra obtained for studied catalyst powders and their CLs.

Catalyst	Form	Fe K β XES		Fe K β HERFD XANES		
		K β IAD / arb. units	K β -M ₁ / keV	pre-edge intensity / arb. units	WL intensity / arb. units	1 st inflection point / keV
Fe0.5	powder	0.287	7058.8	2.57	21.10	7110.9
	CL: NCR* = 0	0.308	7058.9	1.62	22.87	7111.5
	CL: NCR = 0.25	0.335	7059.1	1.62	23.05	7112.3
	CL: NCR = 0.50	0.335	7059.0	1.37	23.71	7112.8
	CL: NCR = 1.00	0.341	7059.1	1.33	24.32	7113.0
	CL: NCR = 2.00	0.337	7059.0	1.35	24.46	7113.3
C-PPy	powder	0.284	7058.9	2.45	20.48	7110.1
	CL: NCR = 0.23	0.309	7058.9	1.78	22.59	7112.7
Porph/C _{prec}	powder	0.275	7058.7	2.46	21.84	7110.3
	CL: NCR = 0.23	0.302	7058.9	2.27	22.77	7110.4

*NCR = Nafion-to-catalyst ratio

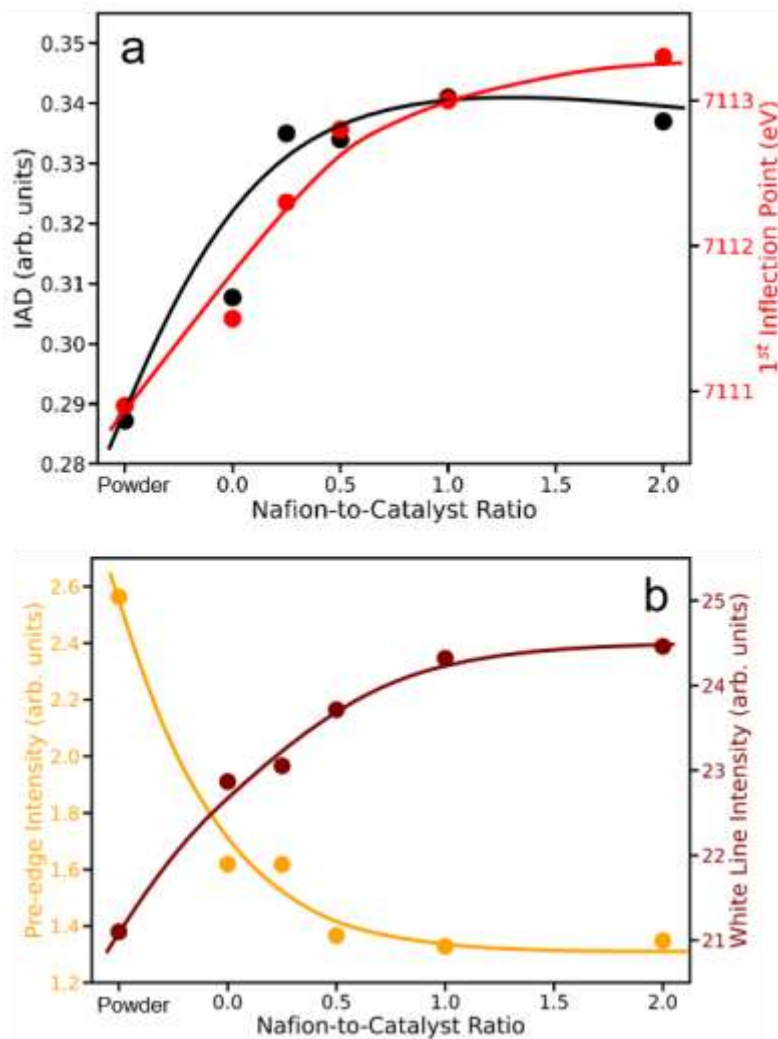


Figure 4. Dependence of Fe K β XES (a) and XANES (a, b) spectral parameters on NCR ratio for Fe_{0.5} catalyst and its electrodes. The lines are provided to guide the reader through the data points and do not possess any physical meaning.

The presence of the solvents (water and alcohols) as well as Nafion binder in the catalyst ink is essential for producing a homogeneous and performing CL. In this work, we demonstrated the particularity of the Fe-N-C catalysts, where the FeN_x sites structure partially evolves from square planar (or square pyramidal) to octahedral by adding $-O_2$ / $-OH$ bond(s) during the electrode preparation procedure and does not fully resemble the one observed in a pristine catalyst powder. While this evolution is universal for Fe-N-C materials, the magnitude of change varies

with the synthesis procedure and consequently with the initial composition of Fe-N-C. The effect is enhanced by Nafion presence in the catalyst ink. This aspect should be taken into account when discussing the catalyst structure – activity – stability/durability correlation(s) as well as analyzing *in situ/operando* measurements performed on these materials, where one should consider the electronic state of Fe in the prepared electrode and not in the catalyst powder.

ASSOCIATED CONTENT

Supporting Information. Details on the XAS/XES measurements and data analysis; catalysis synthesis; CL preparation procedure; additional XANES and XES data.

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Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

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Notes

Any additional relevant notes should be placed here.

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ABBREVIATIONS

CL, catalyst layer; CN, coordination number; GC, glassy carbon; GDL, gas diffusion layer; IAD, integrated absolute difference; MS, Mössbauer spectroscopy; MOF, metal-organic framework; NCR, Nafion-to-catalyst ratio; RT, room temperature; RDE, rotating disc electrode; VtC, valence-to-core; WL, white line; XAS, X-ray absorption spectroscopy; HERFD XANES, high energy resolution fluorescence detected X-ray absorption near edge structure; XES, X-ray emission spectroscopy.

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