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Syntheses of N₂-bridged heterobimetallic complexes, their structural and qualitative bonding analyses

Quentin Le Dé,^a Fourat Orbay,^a Laure Vendier,^a and Antoine Simonneau^{*a}

^a LCC-CNRS, Université de Toulouse, CNRS, UPS, 205 route de Narbonne, BP44099, F-31077 Toulouse cedex 4, France.

antoine.simonneau@lcc-toulouse.fr

Abstract

A series of N₂-bridged ($\mu\text{-}\eta^1\text{:}\eta^1\text{-N}_2$) heterobimetallic complexes were prepared by the reaction of the dinitrogen complex [Mo(depe)₂(N₂)₂] (depe = 1,2-bis(diethylphosphino)ethane) with the group IV-VI high valent metal halides [Cp₂ZrCl₂], NbCl₅, TaCl₅, MoCl₅ and WCl₄ via chloride substitution. Compounds [MoCl(depe)₂($\mu\text{-N}_2$)ZrCp₂Cl] (**1**) and [MoCl(depe)₂($\mu\text{-N}_2$)M'Cl₄(THF)] (**2–5**, M' = Nb, Ta, Mo or W) were crystallized from the reaction mixtures. In the case of the reaction with WCl₄, disproportionation of W(IV) was suspected but not evidenced. Single crystal X-ray diffraction analyses on all compounds allowed a systematic structural study, and the sketching of qualitative molecular orbital diagrams helped discussing the bonding situation and bolster the structural data.

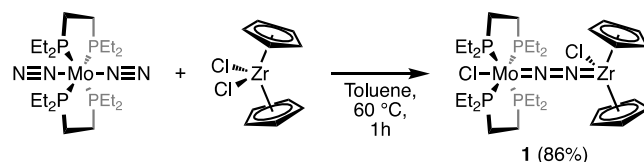
Keywords: Dinitrogen • Heterobimetallic Complexes • Molecular Orbitals • Molybdenum

Introduction

Coordination of dinitrogen to a transition metal center is key to render this inert small molecule reactive. This is the way Nature has elected to achieve N₂ reduction into ammonia in microorganisms expressing the nitrogenase enzymes.¹ Since the isolation of the first N₂ complex [Ru(NH₃)₅(N₂)]²⁺ by Allen and Senoff in the 1960s,² a plethora of coordination and organometallic compounds featuring a N₂ ligand have been described.³ A wealth of coordination modes and reactivity for the dinitrogen ligand is found in those species, depending on the metal, the ligand sphere and the level of reduction of the small molecule, from dinitrogen to diazenide to hydrazide. Reduced N₂ complexes such as diazenido or hydrazido ones are often generated through homobimetallic activation, affording N₂-bridged complexes in either end-on ($\mu\text{-}\eta^1\text{:}\eta^1$) or side-on ($\mu\text{-}\eta^2\text{:}\eta^2$) coordination mode, and are particularly represented among groups III–VI and the lanthanides.⁴ Some of them have allowed the synthesis of hydrazine derivatives by subsequent treatment with electrophiles.⁵ Heterobimetallic N₂ complexes remain scarce, and consequently their reactivity has not been explored,⁶ even though methods to prepare N₂-bridged heterobimetallic

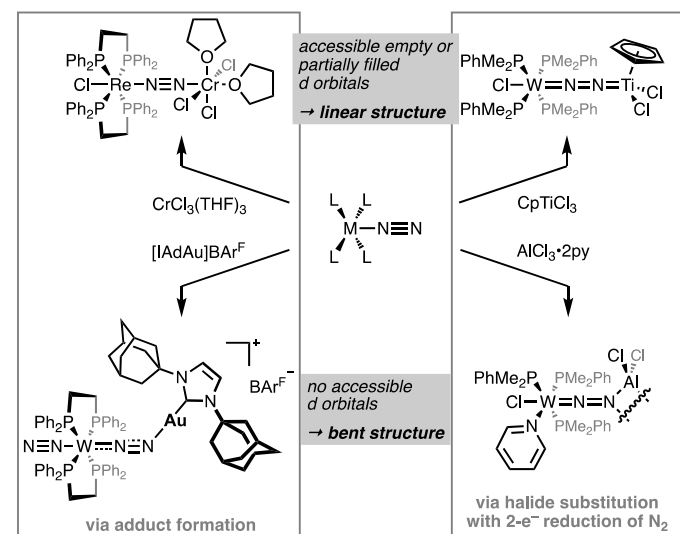
complexes have been long known.⁷ The group of Hidai has studied the reactivity of group 6 dinitrogen complexes with electrophilic metal halides and have obtained diazenido-bridged heterobimetallic complexes via halide substitution. Depending on the presence of accessible and empty d orbital on the electrophile, bent or linear M-N₂-M' array were observed (Scheme 1).^{7f,g,i,k} Contrastingly, for main-group or late metal Lewis acid adducts of dinitrogen complexes^{7c,f,l-q} or those formed from the less electron rich *trans*-[ReCl(L)₄(N₂)] complexes (L = phosphine) with high-valent metal halides, as studied by the group of Chatt,^{7a,b,d,e} reduction of the N₂ unit is usually not observed. Within a research line aimed at exploring the functionalization of dinitrogen within heterobimetallic complexes, we have looked at the reactivity of [Mo(depe)₂(N₂)₂] (depe = 1,2-bis(diethylphosphino)ethane) with group IV-VI high valent metal halides. The goal was to verify whether similar reactivity to those reported by the group of Hidai could be achieved when the more electron-donating depe ligand supports the zero-valent metal and when oxidizing metal halides are employed. We document herein on the synthesis and structural characterization of a series of [Mo]-($\mu\text{-N}_2$)-M'Cl₄(THF) (M' = Nb, Ta, Mo, W) allowing for the systematic structural comparison of heterobimetallic

complexes having identical ligand spheres, plus the isolation of a group IV-VI [Mo]-(μ -N₂)-ZrClCp₂ heterobimetallic compound.



Scheme 2. Synthesis of the Mo-Zr heterobimetallic N₂ complex **1**.

When next turned our attention towards group V chloride salts. The group of Hidai also reported chloride substitution at d⁰ group V CpMCl₄ (M = Nb, Ta) complexes using [W(PMe₂Ph)₄(N₂)₂] or [W(dppe)₂(N₂)₂].^{7j} Equimolar reactions between [Mo(depe)₂(N₂)₂] and NbCl₅ or TaCl₅ were performed in THF, and were characterized by immediate N₂ evolution upon mixing solutions of the reactants. Clean formation of new species was observed after 5 min according to ³¹P NMR monitoring. The product of the reactions of [Mo(depe)₂(N₂)₂] with NbCl₅ and TaCl₅ were characterized by a singlet at 43.8 ppm and 43.6 ppm, respectively. After concentration under vacuum and layering with pentane, orange crystals grew over a night and could be analyzed by sc-XRD. Expectedly, the new complexes [MoCl(depe)₂(μ -N₂)NbCl₄(THF)] (**2**) and [MoCl(depe)₂(μ -N₂)TaCl₄(THF)] (**3**, Scheme 3) were the result of chloride substitution on NbCl₅ and TaCl₅, respectively. Finally, the reactions between [Mo(depe)₂(N₂)₂] and group 6 metal chlorides, MoCl₅ and WCl₄, were performed in THF at room temperature. As in the preparation of **2** and **3**, N₂ evolution was noticed immediately upon mixing solutions of the reactants and after 5 min, ³¹P-NMR analysis of an aliquot no longer showed the resonance associated to [Mo(depe)₂(N₂)₂]. As anticipated due to the paramagnetism of MoCl₅, no new signal in ³¹P NMR could be detected. Intriguingly, the reaction with diamagnetic WCl₄ also led to the disappearance of the NMR signals of [Mo(depe)₂(N₂)₂] without giving new ones. The mixtures were concentrated under vacuum and layered with pentane. Under identical crystallization conditions than for **2** and **3**, crystals suitable for sc-XRD analysis were obtained. Complex [MoCl(depe)₂(μ -N₂)MoCl₄(THF)] (**4**, Figure 1) crystallized in the space group *P*-1. The sc-XRD analysis of the crystalline material obtained from the reaction with WCl₄ was unexpected, since

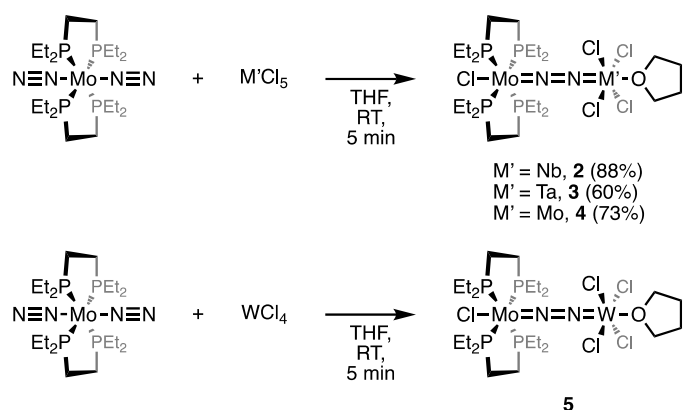


Scheme 1. Examples of preparation of heterobimetallic N₂ complexes from terminal dinitrogen complexes.^{7b,f,j,q}

Results

Our study commenced by treating [Mo(depe)₂(N₂)₂] with the group IV metal complex [Cp₂ZrCl₂]. It was employed by the group of Hidai for the synthesis of heterobimetallic dinitrogen-bridged complexes by halide substitution with *trans*-[W(PMe₂Ph)₄(N₂)₂]^{7g} or [W(dppe)₂(NCS)(N₂)][NBu₄]^{7j} (dppe = 1,2-bis(diphenylphosphino)ethane). A later report from the same research group also documents on similar chemistry employing a series of W(0) dinitrogen complexes and various analogues of [CpTiCl₃],^{7k} but none of these report mentions a productive reaction between [Mo(depe)₂(N₂)₂] and [Cp₂ZrCl₂]. At room temperature in toluene, no reaction was observed according to NMR analysis. After 1 h at 60 °C, the color of the solution changed from orange to purple. In the ³¹P-NMR spectra recorded on an aliquot, the resonance of the [Mo(depe)₂(N₂)₂] complex (56.0 ppm) had vanished and a new signal appeared at 51.3 ppm. After concentration under vacuum and layering with pentane, purple crystals grew over a night and could be analyzed by single crystal X-Ray diffraction (sc-XRD). Complex [MoCl(depe)₂(μ -N₂)ZrCp₂Cl] (**1**, Scheme 2), resulting from chloride substitution on [Cp₂ZrCl₂] by the N₂ ligand of [Mo(depe)₂(N₂)₂], crystallizes in the space group *P*-1 with two molecules per asymmetric unit (Figure 1).

the analogous complex $[\text{MoCl}(\text{depe})_2(\mu\text{-N}_2)\text{WCl}_4(\text{THF})]$ (**5**, Figure 1) was obtained.



Scheme 3. Syntheses of the heterobimetallic complexes $[\text{MoCl}(\text{depe})_2(\mu\text{-N}_2)\text{M}'\text{Cl}_4(\text{THF})]$ ($M' = \text{Nb}, \text{Ta}, \text{Mo}, \text{W}$) **2–5**.

Discussion

Compounds **2–5** have all been obtained in similar conditions, reactions between $[\text{Mo}(\text{depe})_2(\text{N}_2)_2]$ and the metal halide being immediate at room temperature, as witnessed by N_2 evolution right after mixing solutions of the reactants. The high Lewis acidity and unsaturation of the halide salts is likely to explain the quick reactivity with the rather electron rich $[\text{Mo}(\text{depe})_2(\text{N}_2)_2]$ complex. The situation was different for the preparation of **1** as thermal activation was needed. This may result from Cp-induced steric hindrance in $[\text{Cp}_2\text{ZrCl}_2]$ that may lead to a more kinetically demanding ligand interchange. The characterization by sc-XRD of the Mo(II)-W(V) heterobimetallic complex **5** from the reaction with WCl_4 is somewhat surprising and points to a possible disproportionation of W(IV) to W(III) and W(V). Indeed, Cummins and co-workers have reported intermetal chlorine atom transfer between W(IV) and Ti(III), allowing for the synthesis of $[\text{WCl}_3(\text{THF})_3]$.⁸ It may be possible that under our reaction conditions chlorine is transferred from WCl_4 to a transient $[\text{Mo}(\text{depe})_2(\mu\text{-N}_2)\text{WCl}_4(\text{THF})]$ adduct to afford **5** and $[\text{WCl}_3(\text{THF})_3]$ or a related species.⁹ Unfortunately, we were not able to evidence the presence of a reduced W species by sc-XRD by selecting other crystals of the batch, and NMR analysis on a wide spectral window did not reveal the presence of paramagnetic side-products stabilized with THF. Elemental analyses of the crystalline material were

found close to a $5 \cdot \text{WCl}_3 \cdot 1.7 \text{ THF}$ formulation. For all complexes, the N-N stretching band of the complexes could not be unambiguously assigned in the IR spectra due to its shifting to the fingerprint region.

Pertinent bond lengths and angles for compounds **1–5** are compiled in Table 1. The five (hetero)bimetallic complexes show a molybdenum atom supported by two depe ligands within a pseudo-octahedral geometry (Figure 1). A chloride and the bridging N_2 ligand are occupying the apical positions. The structural features of their coordination spheres are almost identical in all complexes. In **1**, the Zr center is found in a distorted tetrahedron. In **2–5**, the high-valent metal lies in a pseudo-octahedral environment with four chlorides in the equatorial plane and one THF molecule facing the bridging N_2 ligand. The ClMoNNM'(O) array slightly distorts from linearity, except in the case of **3**: all 6 atoms were found on a mirror of the $C2/c$ space group with half occupation in the unit cell and were fixed on that mirror plane for refinement, resulting in perfect linearity. The bridging dinitrogen of the complexes have N–N bond lengths between 1.208(7) and 1.262(4) Å. These numbers are pointing to a bond order of two, although in **1**, the recorded N–N bond length is smaller than those found in representative diazene compounds (1.23–1.25 Å).¹⁰ The N_2 unit would act as a doubly reduced dinitrogen moiety, or diazene-diide (N_2^{2-}), with each nitrogen acting as an LX-type ligand for each metal.¹¹ This would result in formal oxidation states of +II for the phosphine-supported molybdenum center and +IV for Zr in **1** and +V for M' in **2–5**. The $(\text{depe})_2\text{Mo-N}$ bond lengths all approach 1.80 Å, the shortest being found in **3** (1.778(3) Å) and the longest in **1** (1.820(6) Å). The Zr–N bond in **1** is 2.009(4) Å and is comparable to the one of $[\text{W}(\text{PMe}_2\text{Ph})_3(\text{py})(\mu\text{-N}_2)\text{ZrCp}_2\text{Cl}]$ ^{7g} (1.98(1) Å, py = pyridine). The $M'\text{-N}$ bonds in **2** and **3** are 1.834(4) and 1.822(3) Å, respectively. For comparison, these bond lengths are longer than what is typically found in group V $[\text{M}(+\text{V})\text{Cl}_3(\text{ether})_n(\text{NR})]$ imido complexes (average of M–N bond lengths from a 13 hits list is 1.758 Å).¹² The same trend is observed for compounds **4** and **5** that show close $M'\text{-N}$ distances of 1.786(5) and 1.784(3) Å, respectively. An average M–N distance calculated for a series of 11 group VI *trans-*

[M(+VI)Cl₄(THF)(NR)] imido complexes is 1.722 Å.¹³ These discrepancies may arise from the nature of the N-based ligand (X₂ vs. LX), the not strictly comparable ligand sphere for the group V halides,

and the difference in oxidation states considered for the group VI halides. Overall, the bond lengths of the Mo-N-N-M' array all agree with double bonding between the four atoms.¹¹

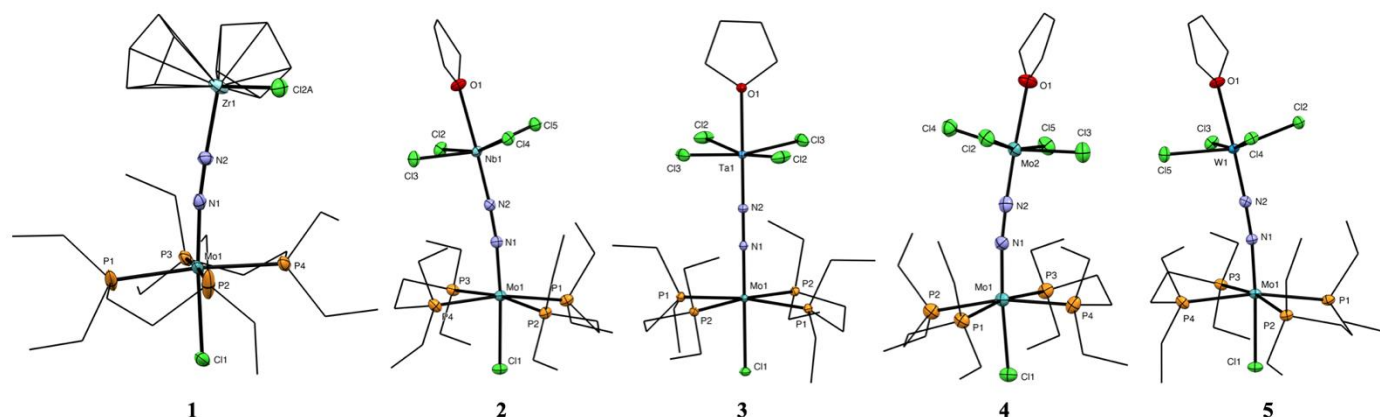


Figure 1. Solid-state molecular structure of compounds **1–5** determined by *sc*-XRD with ellipsoids drawn at the 30% probability level. Hydrogen atoms were omitted and hydrocarbon rings and chains are depicted as wireframes for clarity.

Table 1. Selected bond lengths and angles for compounds **1–5**. n. a. = not applicable.

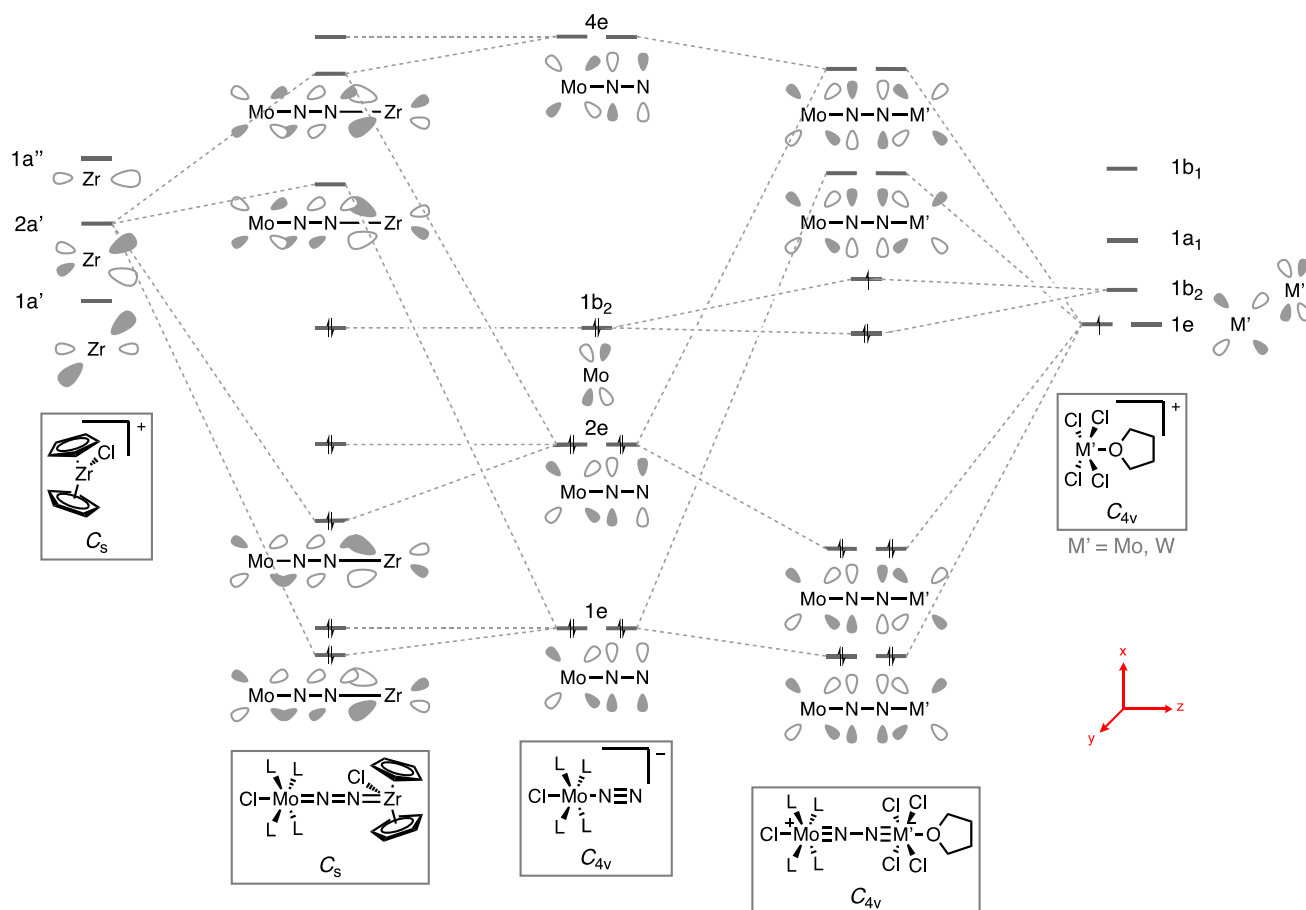
Bond (Å) or angle (°)	1	2	3	4	5
Mo–N	1.820(6)	1.803(4)	1.778(3)	1.816(4)	1.801(3)
Mo–Cl	2.550(2)	2.478(1)	2.4779(9)	2.4517(12)	2.4708(8)
N–N	1.208(7)	1.254(5)	1.258(4)	1.257(6)	1.262(4)
M'–N	2.009(4)	1.834(4)	1.822(3)	1.786(5)	1.784(3)
M'–O	n. a.	2.307(3)	2.254(3)	2.230(4)	2.243(2)
Cl–Mo–N	174.7(2)	173.7(1)	180.0	173.77(12)	173.63(8)
Mo–N–N	173.9(5)	172.8(3)	180.0	173.0(3)	172.4(2)
N–N–M'	176.8(4)	176.6(3)	180.0	177.0(4)	177.4(2)
N–M'–O	n. a.	178.0(1)	180.0	177.91(17)	178.1(1)

The difference between complexes **2–5** is the number of d electrons of the metal opposing the Mo(II) center: the group V–group VI bimetallic complexes **2** and **3** incorporate a MCl₄ moiety that carries no d electrons, while in the group VI–group VI ones (**4** and **5**), the MCl₄ center has a d¹ electronic configuration. Since the N–N bond lengths are comparable in these four complexes, the single electron in **4** should be located on a non-bonding orbital centered on the more oxidized metal and not involved in antibonding orbitals of the N₂ unit or of the M–N linkage.

To shed light on the bonding situation, a qualitative Molecular Orbital (MO) diagram focusing on the π system of the MoNNM' core was drawn from the perturbation of a hypothetical, negatively charged C_{4v} [MoCl(L)₄(N₂)][−] complex by a [M'Cl₄(THF)]⁺ cation of the same symmetry (Scheme 4, right side).¹⁴ The former carries 8 π electrons stemming from the filled {d_{xz}, d_{yz}} set and the two π bonds of N₂, while the latter has either 0 (M' = Nb, Ta) or 1 (M' = Mo, W) d electron in a 1e orbital. In the cationic species, the d_{xy}, d_{xz} and d_{yz} orbital were given high energy levels due to the

destabilizing effects of the chlorides, that are more pronounced for the b_2 (d_{xy}) orbital. Interaction of the frontier orbitals of the two fragments gives a new set of MOs within a heterobimetallic complex of C_{4v} symmetry, in which the single electron of the d^1 cation resides, as suggested by comparison of the crystal structures of **2–5**, in a non-bonding, M' -centered d orbital. A set of two degenerate MOs that are antibonding regarding the N–N linkage are further stabilized than in $[MoCl(L)_4(N_2)]^-$, giving it more N_2 character. Since they are populated, a high level of activation of N_2 (i. e., N–N bond elongation) results. Based on the metrical data issued from sc-XRD analyses and by comparison with computed covalent radii,¹¹ an $Mo=N=N=M'$ depiction for complexes **2–5** makes sense, as mentioned above. However, the drawing of accurate Lewis structures for N_2 -bridged

bimetallic complexes has recently been discussed by Hasanayn, Holland, Goldman and Miller,¹⁵ arguing that comparison of N–N bond lengths with either N_2 , diazene or hydrazine is not always valid. According to their model, compounds **2–5** would better correspond to a $Mo^+(IV)\equiv N-N\equiv M'^-(V)$ depiction, with a positive charge on Mo and a negative one on M' and the N_2 unit acting as a two-fold LX_2 ligand. The former would have 18 valence electrons (VE) and the latter, depending whether it is a group V or VI metal, 16 or 17 VE, respectively, to be compared with 18 VE for Mo and 14 or 15 VE for M' in a $Mo(II)=N=N=M'(V)$ bonding situation (Chart 1). The former depiction better distributes valence electron to the metals and takes into account the electron-withdrawing nature of the $M'Cl_4(THF)$ fragment.



Scheme 4. Qualitative MO diagrams for compounds **1**, **4** and **5** obtained by perturbation of $[MoCl(L)_4(N_2)]^-$ with group IV–VI cations.

Complex **1** has a N–N bond length shorter than the four other complexes (1.207(6) or 1.230(7) Å vs 1.254–1.262 Å). N–N bond lengths are 1.207(6) and 1.230(7) Å, slightly shorter than a N–N double bond (1.25 Å). Zr–N bond lengths are 2.009(4) and

1.991(5) Å long and Mo–N bonds measure 1.821(5) and 1.806(4) Å. Mo–N–N and N–N–Zr angles are 173.96 and 173.35°, respectively. The solid state structure of **1** is very similar to the one of the N_2 -bridged heterobimetallic complex

[WI(PMe₂Ph)₃(py)(μ-N₂)ZrCp₂Cl], reported by Hidai and co-workers^{7g} (N–N and Zr–N bond lengths: 1.24(2) and 1.98(1) Å, respectively). In **1**, the zirconium center has a formal oxidation state of +IV and, like the Ta and Nb ions in complexes **2** and **3**, has a d⁰ electronic configuration. We drew the MO diagram of **1** by the perturbation of the d⁶ complex [MoCl(L)₄(N₂)][−] with the [Cp₂ZrCl]⁺ cation of C_s symmetry to see whether it could provide an explanation for a shorter N–N bond (Scheme 4, left side). The frontier MOs of the latter have been calculated by Lauher and Hoffmann, and we considered here, to match with the orientation of [MoCl(L)₄(N₂)][−], that the Zr–Cl bond is along the y axis.¹⁶ Only the 2a' MO of [Cp₂ZrCl]⁺ can interact with the π system of [MoCl(L)₄(N₂)][−], resulting in the stabilization of only the half of the MOs that are antibonding with respect to the N–N bond. This may explain why the level of activation in **1** is inferior to those in **2–5**. The MO diagram for **1** would support, according to the model of Miller and colleagues,¹⁵ a Mo(IV)=N=N=Zr(IV) depiction. A Mo⁺(IV)≡N–N≡Zr[−](IV) bonding situation similar to the one encountered in **2–5**, where the electron-withdrawing nature of the chloride ligands is helping the stabilization of the negative charge on M', would be counter-intuitive in the case of **1**, the valence electron count for Zr being 20 (Chart 1). The fact that the negative charge would be carried by the more electropositive metal, the scarcity of 20-electron zirconocene complexes¹⁷ as well as the relatively short N–N bond are additional arguments that rule out a hydrazine tetraide (N₂^{4−}) ligand in **1**.

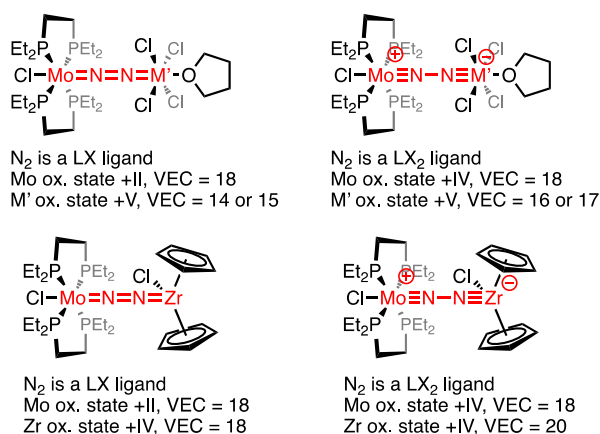


Chart 1. Comparison of bonding situations for the Mo–N–M' core of the heterobimetallic complexes in terms of oxidation states and valence electron count (VEC).

Conclusion

We have prepared new N₂-bridged heterobimetallic complexes, obtained by the reaction of [Mo(depe)₂(N₂)₂] with various group IV–VI halide complexes. Compound [MoCl(depe)₂(μ-N₂)ZrCp₂Cl] (**1**) and a series of analogous [MoCl(depe)₂(μ-N₂)M'Cl₄(THF)] (**2–5**, M' = Nb, Ta, Mo, W) complexes could be structurally characterized. Sketching of the qualitative molecular orbital diagrams helped understand the bonding situation as well as structural similarities or differences. In light of a recent discussion on qualitative analysis of the bonding in bimetallic N₂ complexes by Miller and colleagues,¹⁵ we have applied their model on the compounds reported herein: if a hydrazine tetraide (N₂^{4−}) ligand can be supported for the **2–5** family, a Mo=N=N=Zr depiction (with a diazene-diide ligand) seems more appropriate for **1**.

Materials and methods

1. General information

All reactions were performed in oven-dried glassware with rigorous exclusion of air and moisture, using a nitrogen filled *Jacomex* glove box (O₂ < 2 ppm, H₂O < 1 ppm) equipped with a vacuum line. Solvent used were pre-dried (toluene, tetrahydrofuran and pentane by passing through a Puresolv MD 7 solvent purification machine) degassed by freeze-pump-thaw cycles, dried over molecular sieves and stored in the glove box. THF-*d*₈ (purchased from *Eurisotop*) was degassed by freeze-pump-thaw cycles, dried over molecular sieves and stored in the glove box. MoCl₅, NbCl₅, TaCl₅ and [Cp₂ZrCl₂] have been purchased from *Strem*, the latter one was recrystallized from dichloromethane. WCl₄ has been purchased from *Sigma-Aldrich*. *trans*-[Mo(depe)₂(N₂)₂] has been prepared from MoCl₅.¹⁸ ¹H, ¹³C and ³¹P NMR spectra were recorded in THF-*d*₈ by using NMR tubes equipped with J. Young valves on a Bruker Avance III 400 spectrometer. Chemical shifts are in parts per million (ppm) downfield from tetramethylsilane and are referenced to the residual solvent resonances as the internal standard (C₄H₇D₇O: δ reported = 3.58 and 1.72 ppm for ¹H NMR). ¹³C and ³¹P NMR spectra were calibrated according to the IUPAC

recommendation using a unified chemical shift scale based on the proton resonance of tetramethylsilane as primary reference.¹⁹ Data are reported as follows: chemical shift, multiplicity (s = singlet, t = triplet, q = quartet, quint. = quintet, m = multiplet) and integration. Infrared (IR) spectra were recorded in a nitrogen filled Jacomex glove box ($O_2 < 2$ ppm, $H_2O < 1$ ppm) on an *Agilent Cary 630 FT-IR* spectrophotometer equipped with ATR and are reported in wavenumbers (cm^{-1}). Elemental analyses were performed on samples sealed in tin capsules under N_2 by Analytical Service of the Laboratoire de Chimie de Coordination; results are the average of two independent measurements.

2. Characterization

Synthesis of $[MoCl(depe)_2(\mu-N_2)ZrCp_2Cl]$ (1)

$[Mo(depe)_2(N_2)_2]$ (20.0 mg, 35.4 μ mol, 1 equiv.) was dissolved in 2 mL of toluene. A toluene solution (2 mL) of $[Cp_2ZrCl_2]$ (10.2 mg, 35.4 μ mol, 1 equiv.) was added to the first solution. The mixture was then stirred and warmed to 60 °C during 1 h, time during which the solution became purple. The mixture was concentrated under vacuum to ca. 2 mL and pentane (4 mL) was added. After a night at -40 °C, crystals of (1) were separated from the supernatant and washed with pentane. After drying under vacuum, complex 1 was obtained as purple crystals (25.2 mg, 30.4 μ mol, 86%).

¹H-NMR (400 MHz, THF-*d*₈) δ 6.01 (s, 10H, Cp), 2.01 (m, 8H), 1.92 (m, 8H), 1.79 (m, 8H), 1.23 (t, 12H), 1.14 (t, 12H). ¹³C-NMR (100.6 MHz, THF-*d*₈) δ 110.8 (s), 21.7 (quint.), 18.5 (quint.), 17.9 (quint.), 8.1 (s), 7.5 (s). ³¹P-NMR (162 MHz, THF-*d*₈) δ 51.3 (s).

Elem. Anal. Calcd. for $C_{30}H_{58}Cl_2MoN_2P_4Zr$: C 43.48; H 7.05; N 3.38. Found: C 42.77; H 6.76; N 2.65.

IR (ATR) ν = 2931, 2896, 2875, 1543, 1457, 1407, 1305, 1232, 1182, 1118, 981, 863, 786, 730.

Synthesis of $[MoCl(depe)_2(\mu-N_2)NbCl_4(THF)]$ (2)

$[Mo(depe)_2(N_2)_2]$ (30.0 mg, 53.2 μ mol, 1 equiv.) was dissolved in 3 mL of THF. A THF solution (3 mL) of $NbCl_5$ (14.4 mg, 53.2 μ mol, 1 equiv.) was added to the first solution, and bubbling was immediately

observed. The mixture was stirred for 5 min at room temperature, then concentrated under vacuum to ca. 3 mL. The resulting THF solution was then layered with 4 mL of pentane. After a night at room temperature, crystals of (2) were separated from the supernatant and washed with pentane. After drying under vacuum, complex 2 was obtained as orange crystals (41 mg, 46.7 μ mol, 88 %).

¹H-NMR (400 MHz, THF-*d*₈) δ 2.20 (m, 8H), 2.01 (m, 16H), 1.31 (t, 12H), 1.22 (t, 12H). ¹³C-NMR (100.6 MHz, THF-*d*₈) δ 21.7 (quint.), 20.5 (quint.), 18.1 (quint.), 8.5 (s), 7.4 (s). ³¹P-NMR (162 MHz, THF-*d*₈) δ 43.8 (s).

Elem. Anal. Calcd. for $C_{24}H_{55}Cl_5MoN_2OP_4Nb$: C 32.80; H 6.42; N 3.19. Found: C 32.52; H 6.56; N 2.43.

IR (ATR) ν = 2935, 2874, 1456, 1410, 1352, 1243, 1111, 1031, 869, 837, 817, 759, 734, 673.

Synthesis of $[MoCl(depe)_2(\mu-N_2)TaCl_4(THF)]$ (3)

$[Mo(depe)_2(N_2)_2]$ (30.0 mg, 53.2 μ mol, 1 equiv.) was dissolved in 3 mL of THF. A THF solution (3 mL) of $TaCl_5$ (19.1 mg, 53.2 μ mol, 1 equiv.) was added to the first solution, and bubbling was immediately observed. The mixture was stirred for 5 min at room temperature, then concentrated under vacuum to ca. 3 mL. The resulting THF solution was then layered with 4 mL of pentane. After a night at room temperature, crystals of (3) were separated from the supernatant and washed with pentane. After drying under vacuum, complex 3 was obtained as orange crystals (30.8 mg, 31.9 μ mol, 60 %).

¹H-NMR (400 MHz, THF-*d*₈) δ 2.16 (q, 8H), 1.94 (m, 16H), 1.29 (t, 12H), 1.21 (t, 12H). ¹³C-NMR (100.6 MHz, THF-*d*₈) δ 21.1 (quint.), 20.0 (quint.), 17.7 (quint.), 8.1 (s), 7.0 (s). ³¹P-NMR (162 MHz, THF-*d*₈) δ 43.6 (s).

Elem. Anal. Calcd. for $C_{24}H_{55}Cl_5MoN_2OP_4Ta$: C 29.85; H 5.74; N 2.90. Found: C 30.07; H 5.66; N 2.40.

IR (ATR) ν = 2969, 2937, 2907, 1879, 1602, 1577, 1495, 1480, 1454, 1409, 1379, 1252, 1129, 1082, 1029, 1003, 982, 946, 896, 867, 842, 813, 785, 729, 695.

Synthesis of $[MoCl(depe)_2(\mu-N_2)MoCl_4(THF)]$ (4)

[Mo(depe)₂(N₂)₂] (30.0 mg, 53.2 μmol, 1 equiv.) was dissolved in 3 mL of THF. A THF solution (3 mL) of MoCl₅ (14.5 mg, 53.2 μmol, 1 equiv.) was added to the first solution, and bubbling was immediately observed. The mixture was stirred for 5 min at room temperature, then concentrated under vacuum to ca. 3 mL. The resulting THF solution was then layered with 4 mL of pentane. After a night at room temperature, crystals of (**4**) were separated from the supernatant and washed with pentane. After drying under vacuum, complex **4** was obtained as orange crystals (34.2 mg, 38.8 μmol, 73 %).

Elem. Anal. Calcd. for C₂₄H₅₅Cl₅Mo₂N₂OP₄: C 32.73; H 6.29; N 3.18. Found C 33.12; H 5.96; N 2.92

IR (ATR) ν = 2935, 1875, 1456, 1411, 1376, 1306, 1233, 1183, 1119, 1030, 980, 917, 867, 829, 815, 739.

Synthesis of [MoCl(depe)₂(μ-N₂)WCl₄(THF)] (**5**)

[Mo(depe)₂(N₂)₂] (30.0 mg, 53.2 μmol, 1 equiv.) was dissolved in 3 mL of THF. A THF solution (3 mL) of WCl₄ (17.3 mg, 53.2 μmol, 1 equiv.) was added to the first solution, and bubbling was immediately observed. The mixture was stirred for 5 min at room temperature, then concentrated under vacuum to ca. 3 mL. The resulting THF solution was then layered with 4 mL of pentane. After a night at room temperature, orange crystals were separated from the supernatant and washed with pentane. After drying under vacuum, the crystalline solid weighed 34.2 mg. sc-XRD analysis revealed a W(+V) complex that could not be obtained quantitatively with the applied stoichiometry.

Elem. Anal. Calcd. for C₂₄H₅₅Cl₅MoN₂OP₄W·WCl₃·1.7(C₄H₈O): C 26.76; H 5.03; N 2.03. Found C 26.82; H 4.33; N 1.28.

IR (ATR) ν = 2960, 2933, 2905, 2876, 1894, 1757, 1649, 1454, 1409, 1378, 1308, 1230, 1182, 1121, 1027, 980, 863, 813, 786, 746, 718.

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Declaration of Interests

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

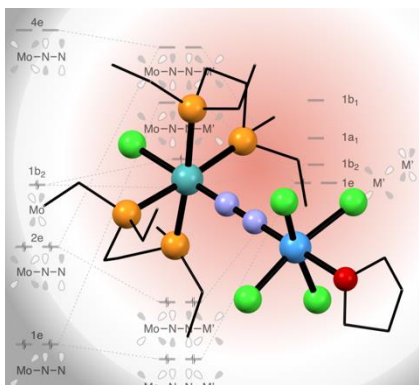
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Graphical Abstract



Synopsis

Syntheses of new N₂-bridged heterobimetallic complexes are reported, with their systematic structural analysis. Similarities or discrepancies are explained with qualitative molecular orbital diagrams.