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Link between medium and long-range order to macroscopic properties of silicate glasses and melts

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Abstract: Understanding the cationic network and the medium range order of glasses and melts allows rationalization of each element's role in terms of the network former, charge compensator or network modifier concepts. Such knowledge can be combined with further details, e.g. regarding the role of highly coordinated network former elements, to model the macroscopic properties of glasses and melts, like their molar volume, density, configurational entropy and viscosity. In this chapter, we present recent advances on such work, which open new avenues for the production of useful and precise models of geologic and industrial glasses and melts.

1. Introduction

The two first chapters of this volume focus on the structure of magmas at different scales: short and medium range order. A quick review of the glass literature, from the 90's or older, reveals that most published papers focused on the structure of glasses, or on properties of glasses or melts. Depending on their original discipline, the work of the scientific community in those areas are different. Physicists published papers on glass structure, including the structure of SiO₂, GeO₂, B₂O₃ as seen by X-Ray or neutron diffraction and Raman spectroscopy for example (Wright 1990, Gaskell et al., 1991, Galeener et al., 1983). Chemists worked on more complex glass compositions, like chalcogenide (Poulain and Lucas, 1970, Poulain 1983), or focused their work on specific properties of glasses like refractive index, density or the oxidation state of multivalent elements (Schreiber, 1986). Earth scientists worked on properties of

melts involved in geologic phenomena, like diffusivity, heat capacity, density or viscosity (Ryan and Blevins, 1987, Carmichael et al. 1977, Bacon 1977; Robie et al., 1979, Bottinga and Weill 1970, 1972). This has changed over time. Since the 2000's, it is easier to simultaneously investigate the structure and properties of glasses. Industrial and technological needs evolved, and the interest of the Earth sciences community became more and more focused on the acquisition and interpretation of *in situ* data to address problems requiring insights about silicate melts at high temperature and high pressure. As a result, the scientific community became more unified, performing more and more studies that simultaneously investigated the structure and properties of glasses and melts. Presently, we have reached a point where it is possible to link together experimental and structural/thermodynamic data to build models for solving industrial or Earth sciences problems. The aim of this chapter is to show how this is possible.

1.1 Glass structure versus macroscopic properties

Among the most important questions about glasses and melts, one is critical for many applications and studies: how the melt/glass structure affects macroscopic properties? For Doremus (1973), a glass is an amorphous solid. The term solid implies a high viscosity, usually greater than 10^{10} Pa.s. This viscosity therefore limits the flow of the body. The amorphous term implies the absence of long-range order (see Figure 5A in the previous chapter, Drewitt et al., 2021), which reveals an analogy with the liquid state. So, a glass is a solid whose properties are similar to those of liquids. Parks and Huffman (1926) even talk about "a fourth state of matter". However, no consensus exists regarding glass, and the nature of glass and the glass transition are two fundamental questions that remain open in condensed matter physics (e.g., see the different definitions in and debate between Zanutto and Mauro, (2017), Popov (2018), Schmelzer and Tropin, (2018)).

An important point is that glasses can exhibit very different chemical compositions and can be prepared by different means. Mineral (i.e. silicate), oxide, non-oxide, metal and organic glasses exist (see for an exhaustive description in Musgraves et al. (2019)). In fact, glasses are found regardless of the type of chemical bonds that link the atoms they contain: covalent, ionic, metallic, Van der Waals or hydrogen bonding. The glassy state is a characteristic of condensed matter. Glasses can be obtained by the rapid quench of liquids, or of a gas by solid condensation, by amorphization of a crystalline phase or by sol-gel methods (Descamps 2017). The most common method for obtaining a glass is by the rapid quench of a liquid¹. During the quench, viscosity continuously increases up to a value so high that the final state can be considered as solid. One can therefore imagine the structure of the glass as being similar to that

¹ rapid cooling rate means 15°/min, the term classic cooling rate is also used.

66 of a liquid whose movements are hindered. Glasses, like undercooled liquids, have a disordered structure
67 that was frozen in at the glass transition temperature.
68

69 A fundamental question is *what happens when a liquid crosses the glass transition temperature?* The
70 latter is the temperature (often abbreviated T_g , and more realistically it is a small temperature interval) at
71 which second order thermodynamic properties like heat capacity show a transition, revealing the locking-
72 in of the atoms in fixed (but undetermined) positions. In essence, T_g is the temperature at which a number
73 of thermodynamic properties go from liquid-like to solid-like values.

74 Unlike a crystallized solid, it is not possible to introduce the term of melting temperature for the glassy
75 state. The solidification of a liquid into glass is accompanied, in fact, by a continuous and gradual increase
76 in viscosity upon cooling, without the appearance of a crystalline structure (Figure 1). This behavior
77 remains true regardless of the glass/melt chemical composition (silicate, aluminosilicate, borate,
78 borosilicate, germanate, chalcogenate, tellurate, metallic, organic... see for details regarding each glass
79 family the [Musgraves et al., 2019](#)). This shows the continuous transition from glass to liquid for a
80 property such as viscosity ([Tamman, 1925](#); see Figure 1). The glass transition is a dynamic phenomenon
81 that characterizes the loss of internal thermodynamic equilibrium. Indeed, the properties of a glass no
82 longer depend solely on pressure and temperature, but also on the temperature at which the glass
83 transition occurs. Finally, Tamman ([1925](#)) sums the concept as “*The viscosity of a liquid increases with
84 increasing undercooling, and in a rather narrow temperature interval it increases very rapidly to values
85 characteristic of solid crystals. A brittle glass is thus formed from an easily mobile liquid. This change
86 in viscosity does not correspond to the behavior of the other properties, which in this temperature interval
87 change relatively only slightly. The change in viscosity is a continuous one and no temperature can be
88 chosen as the freezing-point, the point at which the liquid becomes solid. Glasses are undercooled
89 liquids.*”

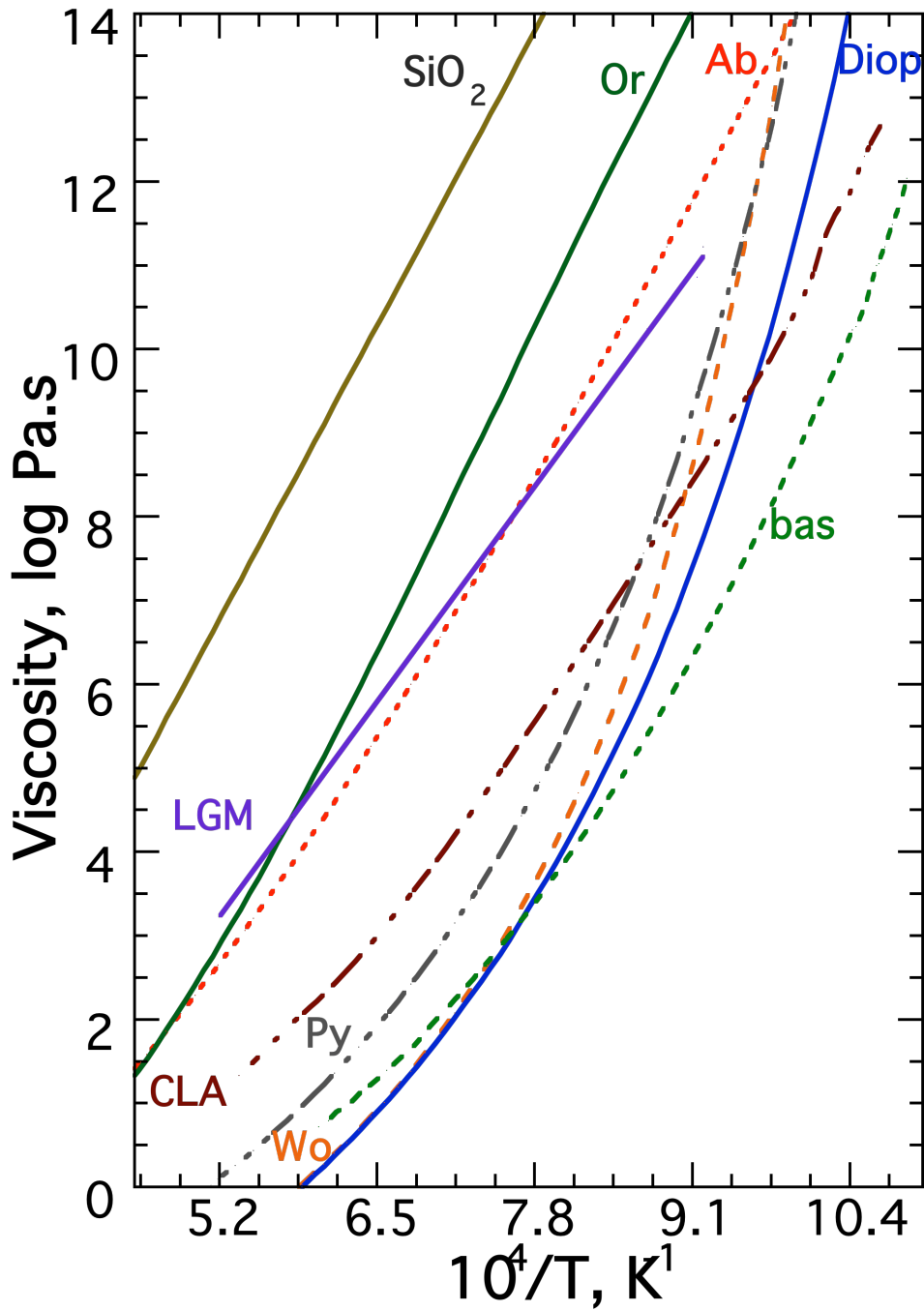


Figure 1: Viscosity versus $1/T$ curves for different melt compositions. LGM and CLA are rhyolite and andesite melts from [Neuville et al. \(1993\)](#), bas is a basalt melt from [Villeneuve et al. \(2008\)](#), SiO_2 is silica melt from [Heterington et al. \(1964\)](#) and [Urbain et al., \(1982\)](#), Ab and Or are albite, $\text{NaAlSi}_3\text{O}_8$, and orthoclase, KAlSi_3O_8 melts from ([Le Losq and Neuville, 2013](#); [Le Losq et al., 2017](#)); Py, Wo, Diop, are Pyrope, $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$, Wollastonite, CaSiO_3 , and Diopside, $(\text{CaMg})\text{SiO}_3$, melt compositions ([Neuville and Richet, 1991](#)).

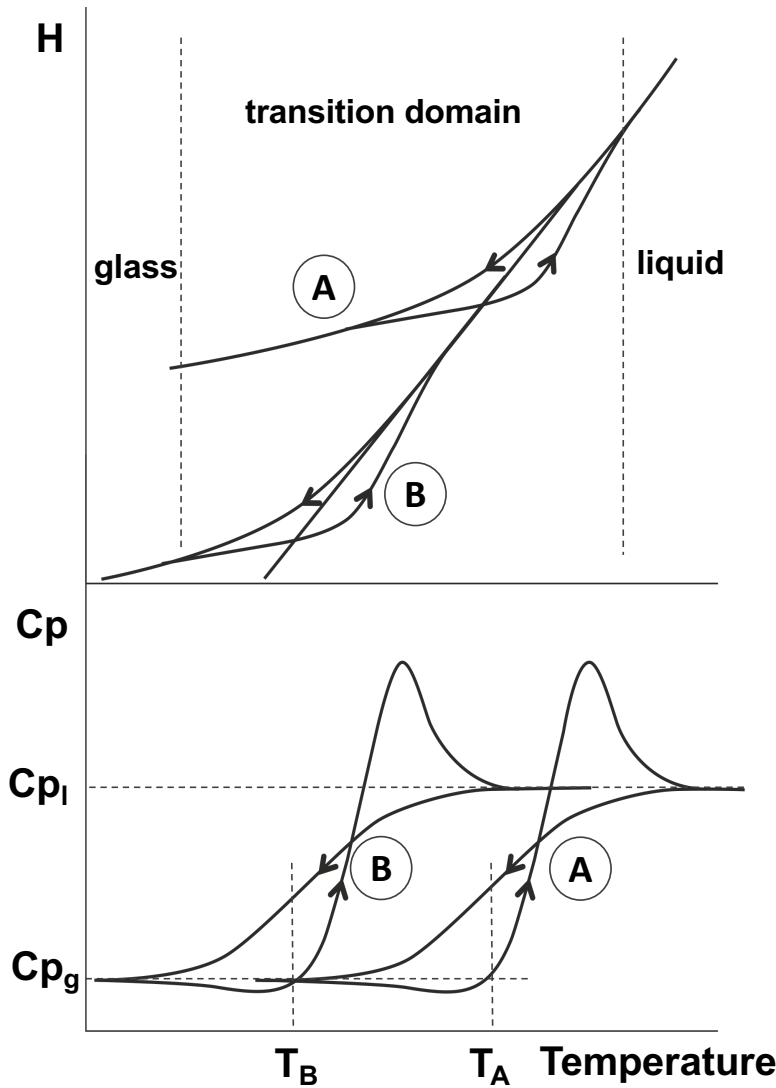
The simplest and earliest characterization of the glass transition is due to Parks and Huffmann (1927), who investigated organic liquids. They note Nernst (1911), has stated that heating a glass “externally it has the properties of a solid, owing to great viscosity and considerable rigidity, produced by strong mutual action of the molecules. An amorphous body differs from a crystal, however, in its complete isotropy and absence of a melting point; on heating, it passes continuously from the amorphous to the usual liquid state, as its properties show steady change with rise of temperature, and no breaks anywhere.”

In fact, Parks and Huffmann (1927) stated “While there is no definite temperature, comparable to the melting point of a crystal, at which all properties undergo a sharp change, there is nevertheless a temperature interval, definite and reproducible, in which a number of properties change with a rapidity approaching that observed in the case of the melting process of a crystal. In brief, there is a softening region instead of a melting point. The glass as it exists below this softening region differs so markedly from the liquid existing above that it might well be considered as a different state of the substance.”

From an energetic point of view, Moynihan et al. (1974) showed that the variations in relative enthalpy and heat capacity with temperature for two different cooling or heating speeds exhibit a sudden change in temperature and corresponds to a dampened variation of these two properties. They showed that for the same liquid, the glass transition temperature varies with the cooling or heating rate. The higher the cooling rate, the higher relative enthalpy or heat capacity (Figure 2). The glass transition thus corresponds to a small temperature and pressure interval upon which properties such as heat capacity undergo a second order transition (Moynihan et al., 1974). This interval is at higher temperatures at high cooling rates (situation A in Figure 2), and moves to lower temperatures at slower cooling rates (situation B in Figure 2).

It is theoretically possible to obtain identical variations for all properties as a function of time at the same temperature (thermal expansion, viscosity for example) but it is not necessary (Moynihan, et al., 1974). With sudden changes in temperature, variations of the liquid properties linked to atomic mobility require some time to reach a new equilibrium. This time is called the relaxation time, τ . A liquid has a large number of different configurational states (Goldstein 1969; 1976). Each state corresponds to a minimum in potential energy and when temperature decreases, the number of possible configurations decreases as the domains for structural rearrangement become larger and larger. As a result, the time for structural relaxation, τ , of the liquid increases. Crossing the glass transition leads to the atoms trapped in given but disorganised positions. This results in quenching the liquid into a glass. The configuration state of the glass does not change from T_g to 0 K, as shown by the fact that the residual entropy of glass remains constant below T_g (e.g., see Tequi et al., 1993; Richet et al., 1991; Richet and Neuville, 1992; Richet et al., 1986; Goldstein, 2011; Schmelzer et al. 2018 and references cited therein). It should be

132 noted that Raman and infrared vibrational spectroscopy data suggest that the structure of a glass is an
 133 image of the instantaneous configuration of the liquid at T_g (Kashio et al, 1980; Kusabiraki and Shiraishi,
 134 1981; Kusabiraki 1986; Sharma et al., 1978; Shevyakov et al., 1978, Neuville and Mysen, 1996).
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136
 137 *Figure 2: relative enthalpy and heat capacity versus temperature for 2 different cooling rates, A and*
 138 *B correspond respectively to fast and low temperature change (redrafted from Moyniham et al., 1974).*
 139 *C_{p_l} and C_{p_g} correspond to the heat capacities of the liquid and the glass, respectively. T_B and T_A*
 140 *correspond to the glass transition at the two different rates A and B, respectively.*
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The uniform cooling of a liquid at a rate $q=dT/dt$ can be likened to a series of instant temperature jumps $\Delta T = T_{fi} - T_i$, where T_{fi} and T_i are the final and initial temperatures, each jump being followed by a Δt period during which the temperature remains constant and equal to T_{fi} . Just after each jump, the viscosity of the liquid, η_l , is still worth η_{Ti} and differs from $\eta_{Tfi} - \eta_{Ti}$ from the new balance value, η_{Tfi} . This is called relaxation, the process by which the liquid tends to reach the state of equilibrium associated with the final temperature T_{fi} . To characterize the kinetics of this evolution, we can define a viscous relaxation time τ_η as equal to:

$$\tau_\eta = \eta_l - \eta_{Tfi} / (\delta\eta/\tau_\eta). \tag{1}$$

Experimentally, τ_η increases when T decreases. Three cases can be distinguished:

$\Delta\tau \gg \tau_\eta$, the substance has a relatively long time to equilibrate its structure at the new temperature. We are in the liquid state or in a glass state at the thermodynamic equilibrium for which the equilibrium viscosity is reached almost instantaneously. During cooling, the viscosity increases to its equilibrium value.

$\Delta\tau \sim \tau_\eta$, this corresponds to the glass transition domain. Viscosity depends on both the time and temperature to which the glass was previously subjected.

$\Delta\tau \ll \tau_\eta$, the relaxation time is much higher than the measurement time. No configurational rearrangement is then possible. The liquid freezes. Only the vibrational part of the heat capacity remains, which is close to the heat capacity of the crystal.

Relaxation times depend heavily on temperature ([Rekhson, 1975](#); [1989](#), [Simmons et al., 1970](#); [1974](#), [Dingwell and Webb, 1990](#)), as well as on the structure and therefore glass composition; e.g. borosilicate glasses have higher relaxation times than silicate or aluminosilicate glass compositions ([Sipp et al., 1997](#)). Figure 3 shows, for a soda-lime silicate glass, that relaxation time increases with decreasing temperature: e.g. equilibrium is reached after 500, 1200, 1600 min respectively at 795, 788 and 777 K.

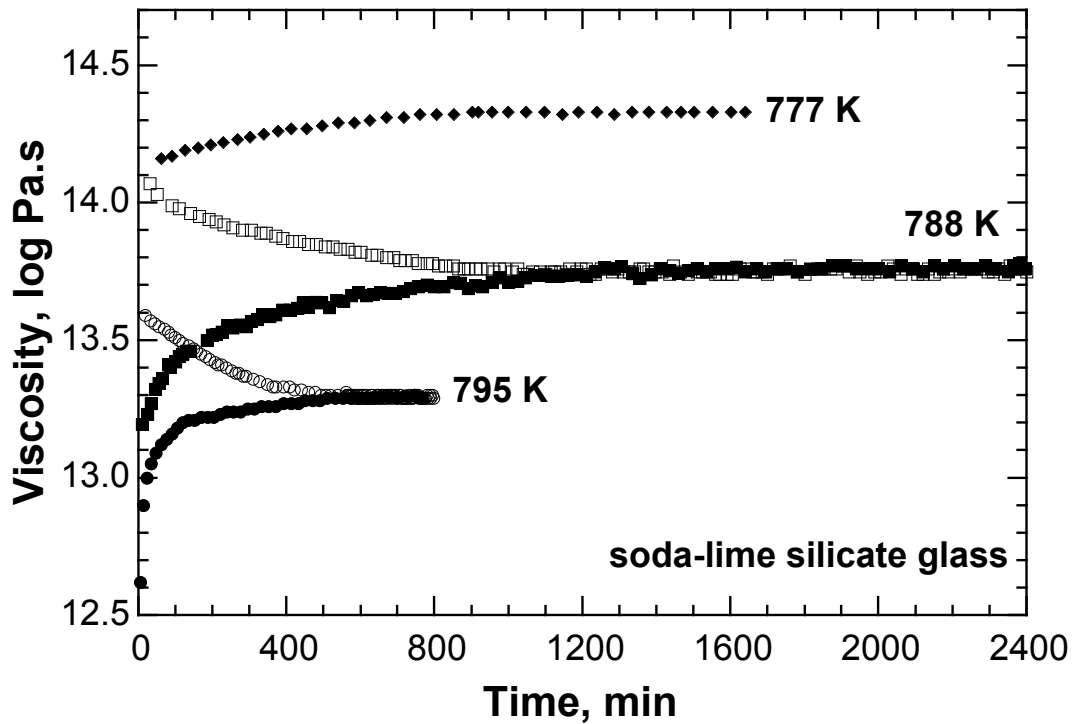


Figure 3: Viscosity versus time for a soda-lime silicate glass at different temperatures (redrafted from Sipp et al., 1997). We can note that increasing viscosities represent relaxation after previous measurements performed at higher temperatures, whereas decreasing viscosities indicate relaxation after annealing at lower temperatures.

However, it is important to note that there is no reason why the relaxation times of the various properties should be identical at equal cooling speeds. Moynihan et al. (1976a) indicate that “for example, the enthalpy changes occurring during the approach to equilibrium of a network glass following a change in temperature could be considered to arise from the breaking of some of the network bonds, while the volume changes are due to rearrangement of the structure into less densely packed configurations. One cannot say a priori, however, which of these two processes should on the average occur earlier in time. That is, bond breaking might be a necessary precursor to volume-changing rearrangements of the structure, so that one would expect H to relax faster than V . On the other hand, configurational rearrangements leading to large volume changes might occur early in time, stressing the network and leading to subsequent bond breaking, so that one would expect V to relax faster than H .” In practice, it appears that the differences in relaxation time between volumetric and calorimetric relaxation times are relatively small, if not almost non-existent, and are essentially due to differences in cooling or heating

rates (Sasabe et al., 1977; Moynihan and Gupta, 1978; Dingwell and Webb, 1990). Differences between relaxation times for the same material but determined with methods sensitive to different properties are clearly visible when comparing data from dilatometry, calorimetry and viscosimetry. For example, in the case of alkaline-earth aluminosilicate glass compositions, drop calorimetry T_g are generally found at temperatures at which viscosities are of $10.7 \pm 0.5 \log \text{ Pa.s}$, a value lower than that of $\sim 12 \log \text{ Pa.s}$ that is known to be the reference for the determination of the viscous T_g (Neuville and Richet 1991; Tequi et al., 1991).

1.2 Thermodynamic approach to glass transition

We have seen that temperature and pressure are not sufficient parameters to characterize the state of glass. Tool and Eichlin (1931) introduced the concept of fictive temperature, T_{fic} , to characterize a glass at constant pressure. It can be defined as the temperature at which the glass would be in an equilibrium state (as a melt) if it could be heated and measured instantly. The fictive temperature can be defined during cooling in dilatometry or calorimetry, as the glassy transition temperature.

It is possible to give a formal thermodynamic definition of the fictive temperature, considering it as an order parameter. Moynihan et al. (1974, 1976a,b) discussed fictive temperatures, their influence on enthalpy properties and order parameters. It can be recalled, that the glass transition looks like a second-order transition for which volume, enthalpy and viscosity are continuous functions of T but not heat capacity, dilatation and compressibility coefficients. For the latter parameters, a steep transition is observed when temperature crosses the glass transition.

In reality, fictive temperature is a critical parameter as it reflects how viscosity, hence relaxation time, can show non-equilibrium behavior and can depend on time in the supercooled domain. In Figure 3, the effect of fictive temperature is clearly visible. The downward curve is obtained from a sample that has a fictive temperature below the measurement temperature. Similar behaviors are visible for density, or refractive index (Winter, 1943, Ritland, 1954). When the fictive temperature of the glass is higher than the measurement temperature, the glass has a more disordered configurational state than the state it should have if the glass were in thermodynamic equilibrium, i.e. its "fictive" configuration entropy is too high. It has a lower viscosity than its equilibrium viscosity, so its viscosity increases over time until the fictive temperature equals the measurement temperature. For a fictive temperature lower than that of measurement, the viscosity decreases until the establishment of the thermal equilibrium characterized by the equality of the two temperatures. In Figure 2, the relaxation curves over time for the cases of cooling

or heating are asymmetrical. This is because a liquid increases its entropy faster with an increase in temperature (Rehson, 1980, Dingwell and Webb, 1990; Sipp et al., 1997)².

1.3 Viscosity and glass transition

The glass transition is closely related to transport phenomena and more specifically to viscosity (Adam and Gibbs 1965; Gibbs and Dimarzio, 1958; Goldstein, 1969; Grest and Cohen, 1980; Scherer 1984, Richet, 1984), one of the fundamental properties of liquids that intervenes in transport processes. For the experimentalist, liquid viscosity also can be considered as a structural probe. Many models attempt to describe the variation of viscosity with temperature. Among them, we can cite the empirical Arrhenius and Tamman-Vogel-Flucher (TVF) equations, or the thermodynamic Adam-Gibbs (AG) model (see also section 2) allows linking melt mobility to its thermodynamic properties such as heat capacity and configurational entropy, and can be used for viscosity predictions and extrapolations (Neuville and Richet, 1990). Le Losq and Neuville (2017) further showed that the AG model allows linking melt structural knowledge to thermodynamic properties and viscosity, in order to build complete, extensive models for property predictions.

1.4 Configurational properties and glass structure

Configuration properties are the key to understand the difference between liquids and glasses. To illustrate the importance of such configurational aspects, consider the second order thermodynamic properties of crystal, glass and liquid pyrope ($\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$).

TABLE 1: Heat capacity of pyrope and $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$, glass and liquid in $\text{J mol}^{-1} \text{K}^{-1}$. Data are from Tequi et al. (1991).

T(K)	Crystal	Glass	Liquid
298	325,6	327,8	
1020	492,2	504,4	665,7
1700	522,6		689,5

Above room temperature, the glass and crystalline phase have similar heat capacities, C_p . On the contrary, the C_p of the liquid is 25% higher than that of solids (crystal or glass). Moreover, it should also be noted that the C_p of the liquid usually increases with temperature of the liquid and cannot be considered as a constant. These differences can have a significant impact when you integrate with

² A fictive pressure, P_{fic} , can be defined as the pressure at which the glass would be in an equilibrium state if it were subjected instantly to pressure equal to P_{fic} , at constant temperature.

temperature or pressure, as shown by the example of pyrope and $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ liquid. For instance, a simple extrapolation of the heat capacity of the $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ liquid to low temperatures would produce a liquid with a lower entropy than that of its crystalline form below 694 K (Figure 4A). This temperature actually corresponds to the Kauzmann temperature, a temperature below which the liquid entropy would become lower than that of a crystal with the same composition. Such a system does not exist because a property of a liquid like the entropy should always remain higher than that of its crystalline form at a given temperature. The violation of this idea is known as the Kauzmann paradox (Kauzmann, 1948)³. The pyrope example shows that configuration terms play an important role on the properties of liquids. It is therefore necessary to understand how configurational properties vary, and their link to melt and glass structure. Figure 4A shows the variation of the entropy of a crystal, glass and liquid for the $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ composition as a function of temperature. It is clear that the residual entropy corresponds to the difference between the entropy of the glass and that of the crystal at 0 K. This residual entropy is very small compared to the entropy of glass, crystal and liquid. This residual entropy is called configurational entropy, S^{conf} , and it is the key to understanding glasses and liquids. To determine S^{conf} , it is necessary to solve the entropic cycle represented in Figure 4A. However, this is only possible for minerals that melt congruently, and with a composition that can exist in a stable form as crystal, liquid and glass. Measuring the heat capacity of the crystal, C_{pc} , from 0 K to the melting temperature, T_m , allows determining its absolute entropy S_c at T_m :

$$S_c = \int_0^{T_m} \frac{C_{pc}}{T} dt. \quad (2)$$

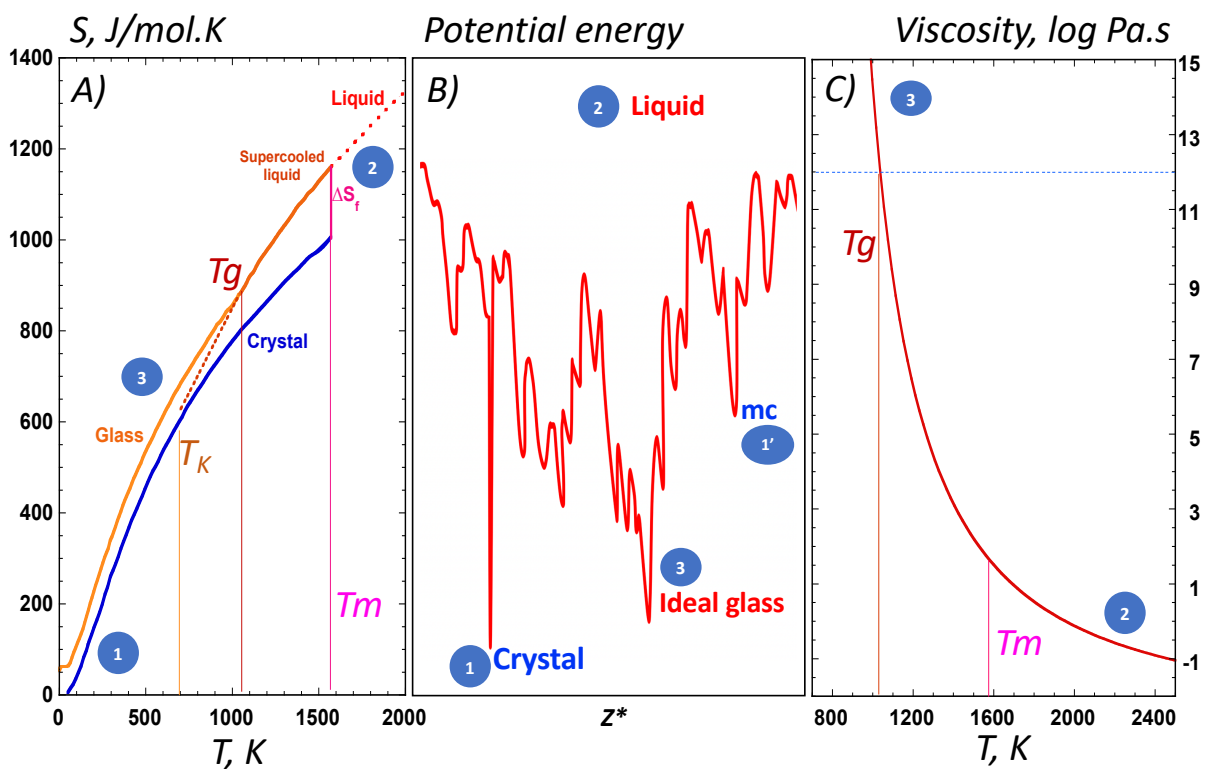
By adding the entropy of fusion of the crystal, ΔS_f , to S_c , the entropy of the liquid S_l is then obtained at T_m . Measuring the heat capacity between the melting temperature, T_m , and the glass transition temperature, T_g , allows determining the heat capacity of the liquid and the supercooled liquid, C_{pl} . The measurement of the heat capacity of the glass between T_g and 0 K provides C_{pg} . Finally, the residual entropy, also known as the so-called configurational entropy S^{conf} , is obtained at 0 K via:

$$S^{conf}(0\text{ K}) = \int_0^{T_m} \frac{C_{pc}}{T} dt + \Delta S_f + \int_{T_m}^{T_g} \frac{C_{pl}}{T} dt + \int_{T_g}^0 \frac{C_{pg}}{T} dt. \quad (3)$$

This calculation is illustrated in Figure 4A for a composition of $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ (Tequi et al., 1991). As this Figure shows, S^{conf} represents a small difference between large numbers, and its value will be affected by a very large error if calorimetric measurements are not made with the highest possible accuracy. For this purpose, the studies focused on determining S^{conf} from calorimetry measurements require high-precision measurements of thermal capacity: heat capacities need to be determined better

³ this well-known paradox in temperature is recently proposed via pressure, but without any experimental basis see papers of Schmelzer et al., (2016) and Mauro (2011).

277 than 0.2 % by adiabatic calorimetry between 0 K and ambient, and better than 0.5 % by high-temperature
 278 drop calorimetry (e.g. [Tequi et al., 1991](#)). Aside from the time and the difficulty of performing such high
 279 precision heat capacity measurements, one of the main drawbacks of the calculation of the
 280 configurational entropy through the thermodynamic cycle is the necessity to study minerals that melt
 281 congruently, with compositions existing as glass, liquid and crystal as stated previously. Only a few
 282 chemical compositions can be studied in this way. Fortunately, as we shall see later, S^{conf} can be
 283 determined through the use of the Adam and Gibbs model combined with viscosity measurements.



284
 285 *Figure 4: A) entropy versus temperature ([Tequi et al., 1991](#)), B) schematic representation of the com-*
 286 *plex higher dimensional potential energy hypersurface of an N -particle system in terms of a two-dimen-*
 287 *sional diagram of chemical potential versus some collective configuration coordinate Z^* , which defies*
 288 *precise definition, but the width of the funnel represents entropy ([Angell, 1991](#); [Mossa et al. 2002](#)),*
 289 *and C) viscosity versus temperature for the $Mg_3Al_2Si_3O_{12}$ composition ([Neuvill and Richet, 1991](#)). The*
 290 *same numbers are all connected together in the three figures. 1 is the Pyrope crystal that presents the*
 291 *minimum potential energy. 1' is a metastable crystal, mc, that can have a potential energy higher than*
 292 *that observed for the glass. 2 is the liquid with a low viscosity, high entropy and high potential energy,*
 293 *in which atoms can move rapidly from one position to one other. 3 is the glass, in which atoms can*
 294 *move very slowly, there is a relaxation time, viscosity increase and entropy decrease, and the atom can*
 295 *fall in different potential energy minima and stay inside or jump to another with a lower potential en-*
 296 *ergy. T_m is melting temperature, T_g , glass transition temperature, T_K Kauzmann temperature, ΔS_f en-*
 297 *tropy of melting at the melting temperature. Supercooled liquids are liquids between melting tempera-*
 298 *ture and glass transition temperature.*

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302 Figure 4B schematically illustrates the positions of atoms in a crystal as a function of the collective
303 configuration coordinate, as determined by local minima of interatomic potentials that are at specific
304 positions, allowing the building of a repetitive pattern characterized by long range order. In a glass, the
305 bond angles and interatomic distances are not constant but extend over a range of relatively close values.
306 Long-range order does not exist, and one can represent the interatomic potentials as a plane showing
307 minima separated by barriers with varying shapes and heights. As a result, glass entropy is higher than
308 that of crystal at given temperature (point 3, Figure 4). The ideal glass with the minimum potential exists,
309 but also other minima can exist with higher potential for the glass state or even the metastable crystal
310 (point 1'). Now suppose that a certain amount of heat is brought instantly to the glass. Below the glass
311 transition, the thermal energy brought upon increasing temperature is accommodated as vibrations, via
312 an increase in average vibration amplitudes. The specific heat is vibrational in nature and the material
313 behaves like a solid. At and above T_g , the thermal energy becomes important enough to allow the atoms
314 to cross the energy barriers that separate the different configurational states (Richet and Neuville, 1992).
315 This configurational contribution is necessarily positive, and the liquid or glass states can be defined by
316 the existence or absence of this contribution (Davis and Jones, 1953). In a way, the glass transition can
317 be seen as the beginning of the exploration by atoms of positions corresponding to the highest values of
318 interatomic potential (Goldstein, 1969). This distribution of configurations over increasingly high
319 potential energy states is the main characteristic of atomic mobility and low viscosity or relaxation time
320 (point 2, Figure 4).

321 If we now look at the changes in volume in an amorphous material, we notice that a general
322 characteristic of interatomic potentials is their anharmonic nature, i.e. that the forces applied to the
323 vibrational atoms are not exactly proportional to the movements in relation to the equilibrium positions
324 of these atoms (Richet and Neuville, 1992). An increase in vibration amplitudes therefore leads to an
325 increase in interatomic distances. Like solids, liquids also have such anharmonic vibrational expansion,
326 but higher-energy patterns that begin to be explored over the glass transition are generally associated
327 with increased interatomic distances. This is why the thermal expansion coefficient generally increases
328 significantly at the glass transition (Richet and Neuville 1992).

329

330 1.5 pressure-temperature space

331 As we have mentioned before, there is a glassy transition pressure (Rosenhauer et al., 1979), which
332 allows one to place the glass transition in a pressure-temperature plane for a given cooling or heating

rate. However, the effects of pressure and temperature on the properties of glasses are actually different. High pressure can produce irreversible configurational changes. So, for given experimental quench rates or measurement time scales, the kinetics of configurational changes are markedly different depending on whether they are caused by pressure or temperature. The main reason for this difference is that the shape of potential energy wells varies little with temperature but significantly with pressure. If high kinetic energy is required to cross the potential barriers at constant pressure, changes in these barriers with pressure can lead to new configurational states at lower temperatures, if the pressure is high enough. The processes occurring at high pressure in glasses are likely similar to those that occur in liquids. Since the 2000's, many papers have shown significant changes in the glass structure with increasing pressure, more or less linked to changes in their properties (e.g. [Poe et al., 2001](#), [Suzuki et al., 2002](#), [Wang et al., 2014](#)). In this volume, an exhaustive chapter can be found on the effect of pressure on the structure and properties of melts and glasses by Sakamaki and Ohtani ([2021](#)).

2. Silicate glasses and melts

2.1 Alkali or earth alkaline silicate glasses and melts

As depicted in the previous two chapters, silica glass is a material well-connected at the molecular scale, built from a three-dimensional network of tetrahedral units composed of covalently bonded central Si and apical O atoms (a.k.a. Bridging Oxygens or BO). However, contrary to α -quartz that forms perfect trigonal crystals, SiO_2 glass is built by SiO_4 tetrahedra distributed disorderly yet not totally randomly (e.g. see Figure 5A in chapter 3, this volume). The relative density d of silica glass equals 2.20 (molar volume $V_m = 27.311 \text{ cm}^3$), a value lower than that observed for α -quartz ($d = 2.65$, $V_m = 22.673 \text{ cm}^3$) that presents a compact and well-ordered structure. This implies that silica glass presents a rather porous structure, with possibly a minor amount of Non-Bridging Oxygen (NBO, see Chapter 2, this volume) atoms not connected to 2 Si atoms ([Brückner 1970](#); [Fanderlik, 1990](#)). The fraction of NBO in silica is usually considered as negligible, making the silica structure a very strong one.

As an alkali oxide is added, Na_2O for example, more and more NBO are created until reaching the possibility to create isolated SiO_4 tetrahedra linked together by only ionic Na-O bonds. For alkali silicates, this corresponds to a theoretical view because Na_4SiO_4 glasses do not exist. However, in the Ba-silicate system, it is possible to obtain glasses with 33% and 37% silica with classic cooling rate ([Bender et al., 2002](#)), while in the Mg-silicate system it is possible to obtain glasses up to 62 MgO in mole percent (38% SiO_2) but below 50 percent of silica, it is necessary to use a fast cooling rate as shown in Raman spectra ([Neuville et al., 2014](#)). Glass formation in the SiO_2 - Na_2O system is actually not possible for amounts of Na_2O lower than that of the metasilicate composition Na_2SiO_3 ([Schairer and Bowen, 1956](#)). At higher Na_2O contents, up to 58 mole percent, it is possible to obtain Na silicate glasses only

via high-speed quenching of a small quantity of melt (a few grams; Imaoka and Yamazaki, 1963). Stable, large pieces of glass cannot be formed below 60 mol% of SiO₂ in this system (Neuvill, 2006). For other alkali-silicate glasses, it is possible to make glasses up to 35 and 55 mole % of Li₂O and K₂O, respectively (Imaoka and Yamazaki, 1963, Levin et al., 1964).

To summarize, alkali silicate glasses can be easily obtained at normal cooling rates in the 60 to 99 mole % SiO₂ range (see Raman spectra of silicate glasses in Neuvill et al., (2014)), and alkaline-earth silicate glasses between 33 and 75 mol% SiO₂, particularly with heavy alkaline-earth elements like Ba. For Mg, the glass forming domain is smaller, and ranges between 37 and 55 mole % SiO₂ (Imaoka and Yamazaki, 1963, Richet et al., 2009). In the CaO-SiO₂ system between 1 and 62 mole % SiO₂, the melts at temperatures just above the liquidus consists of two immiscible phases (Imaoka and Yamazaki, 1963, Levin et al., 1964, Hudon and Baker 2002ab; Neuvill 2006).

Alkali or alkaline-earth silicate glasses show almost linear variations in molar volume with addition of M₂O or M²⁺O in silica glasses (Figure 5). Molar volumes obtained from density measurements for CaO-Al₂O₃-SiO₂ and Na₂O-Al₂O₃-SiO₂ are also available in the literature (Seifert et al., 1982, Doweidar 1998, 1999) but we have chosen not to report them on the figures for two reasons: -i) they show similar trends to ours, - ii) they are made with different cooling rate than ours (not mentioned in the articles). The molar volume (V_m) of silicate glasses increases or decreases as a function of the cation size⁴. In the case of cations presenting an ionic radius larger than O²⁻ (132 pm), like Cs⁺, Rb⁺ or K⁺, the glass molar volume increases with cation addition. For cations presenting an ionic radius lower than that of O²⁻, like Na⁺ and Li⁺, V_m of silicate glasses decreases both with the cation size and its amount. A similar behavior is observed in alkaline-earth silicate glasses. In Ca and Mg silicate glasses, V_m varies almost linearly with the glass SiO₂ (Figure 5). In Figure 5, a dotted straight line is drawn between SiO₂ and MgO with MgO with Mg in 6-fold coordination, and it is clearly observed that the V_m of Mg-silicate glasses are below this line, which suggests that the CN of Mg is significantly less than 6 in amorphous magnesium silicates. Such an interpretation agrees with ²⁵Mg Nuclear Magnetic Resonance (NMR) spectroscopy that shows that Mg is essentially in four-fold coordination in silicate glasses (Fiske and Stebbins, 1994; Georges and Stebbins, 1998, Kroeker and Stebbins, 2000, Shimoda et al., 2007a,b), as well as with XANES at the Mg K-edge (Trcera et al., 2009) while Ca is essentially in six fold coordination in silicate glasses as shown by XANES at the Ca K-edge (Neuvill et al., 2004b, Cormier and Neuvill, 2004; Ispa et al., 2010; Cicconi et al., 2016). These gentle variations may indicate that Ca and Mg remain almost in the same

⁴Cation size are given in Whittaker and Muntus (1970) and molar volumes are from Robie et al. (1979).

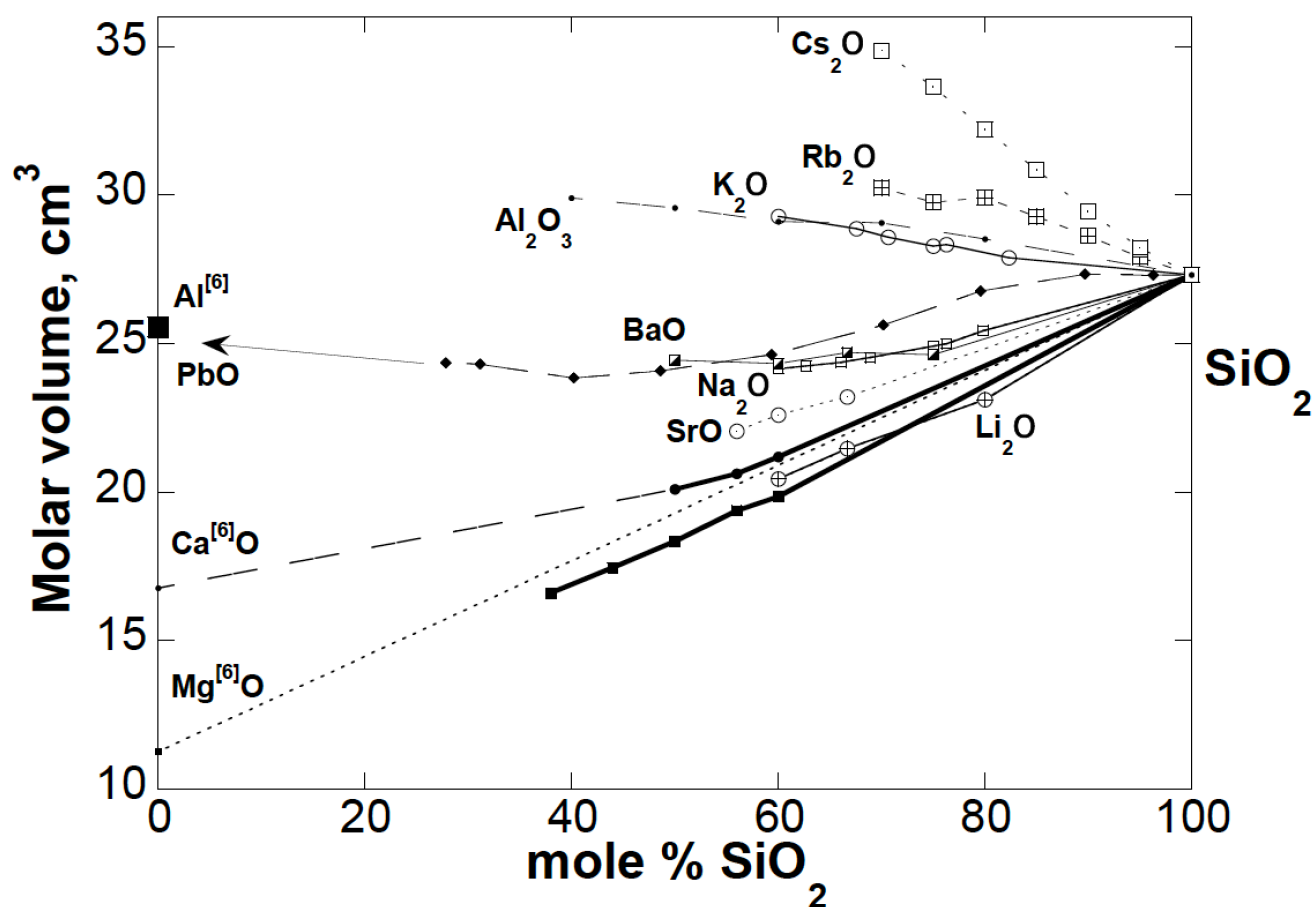


Figure 5: Molar volume of glasses at room temperature: data for Ca, Mg, Na, K, Li silicate glasses are original data from Neuville, Sr and Ba Novikov (2017) Rb, Cs O'Shaughnessy et al., (2017, 2020); Pb-silicate (Ben Kacem et al., 2017) and Al-silicate glasses (Wang et al., 2020). Molar volume of ^[6]Al corresponds to Al₂O₃, and MgO and CaO are from Robie et al. (1979). Density of MA44.00 and MA38.00 glasses are respectively 2,807 and 2,881, Raman spectra of these two glasses are given by Neuville et al. (2014).

Viscosity of silicate melts

In Figure 6A, the viscosity of SiO₂ and soda silicate melts are plotted as a function of reciprocal temperature. Silica shows the highest viscosity of all glass compositions, as well as a strongly linear behavior indicating an Arrhenian behavior (slope is equal to an activation energy).

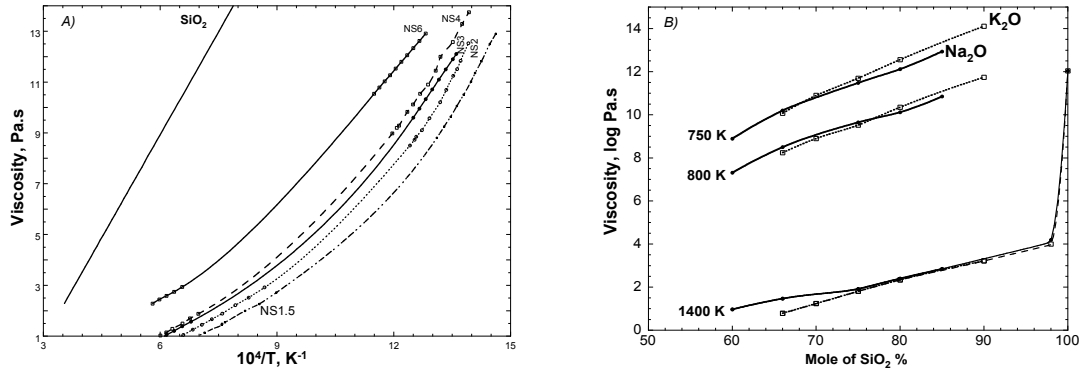


Figure 6: A) viscosity versus reciprocal temperature (SiO_2 NS6, NS4, NS3, NS2, NS1.5, data from [Neville, 2006](#), and [Le Losq et al., 2014](#)). NSx: x corresponds to the ratio of SiO_2/Na_2O in mole percent (e.g. NS4: $x = 4 = 80/20$ - 80 % SiO_2 and 20 % Na_2O); B) viscosity versus SiO_2 content for Na- and K-silicate glasses (data from [Poole, 1948](#), [Bockris et al., 1955](#), [Neville, 1992](#), [2006](#)).

The fact that silica easily forms a glass, albeit a strong one with a very high viscosity, directly indicates that Si is a network former cation. Other elements, presenting lower valence and larger ionic radius, can play different roles in the glass structure. For instance, the introduction of Na_2O in silica glass yields to breaking the covalent Si-O-Si bridges, and, hence, to the formation of Non-Bridging Oxygens, NBO. Na is considered, in this case, as a network modifier cation. Na fits into the glass structure as cations linked to the surrounding oxygens by bonds that are much more ionic and thus weaker than Si-O covalent bonds. Thus, the structure of sodium silicate glasses is weaker than that of vitreous silica. This translates into a lowering of the viscosity at given temperature when adding Na_2O in SiO_2 (Figure 6A). Furthermore, the introduction of Na_2O in silica glass produces a non-Arrhenian behavior that increases with x for the NSx glass; x corresponds to the ratio of SiO_2/Na_2O in mole percent (Figure 5A). When viscosity varies with $1/T$ following an Arrhenian behavior, it follows this equation:

$$\text{Log } \eta = A + B/T, \quad (4)$$

with A the viscosity at infinite temperature and B an activation energy. This equation can only be used for a few chemical compositions like SiO_2 , GeO_2 , $NaAlSi_3O_8$ or $KAlSi_3O_8$. An Arrhenian behavior implies that the activation energy, B, is independent of temperature, and, hence, that the process underlying melt viscous flow remains the same regardless of temperature. For melts showing a strong non-Arrhenian behavior like NS3, the activation energy term, B, decreases from 2000 kJ mol⁻¹ at 1000 K down to 300 kJ mol⁻¹ at 1800 K. In such conditions, it is no longer possible to use the Arrhenian

equation. For this reason, Tamman, Vogel and Fulcher proposed the so-called TVF equation (Vogel, 1921; Fulcher, 1925, Tamman and Hesse, 1926) to link viscosity and temperature:

$$\text{Log } \eta = A + B/(T-T_l), \quad (5)$$

A , B and T_l are just fitting parameters without physical meaning. The TVF equation is very useful to fit and interpolate viscosity data at low and/or at high temperature (Neuville and Richet, 1990). This equation has been used in many different papers and is at the basis of several viscosity models (Bottinga and Weill, 1971; Shaw, 1972; Persikov 1991; Giordano et al., 2008; Giordano et Russell, 2018). The empirical parameters A , B , T_l can be linked to glass structure (Giordano and Russell, 2018). However, while finding very practical applications, this equation remains purely empirical.

Aside from the empirical Arrhenius or TVF equations, the Adam-Gibbs equation, derived from the theory of Adam and Gibbs (1965), is one of the best candidates to fit and extrapolate viscosity measurements for silicate melts (Urbain, 1972, Wong and Angel, 1976; Scherer, 1984, Richet 1984, Neuville and Richet, 1990; 1991, Mauro et al., 2009). The strength of this equation is that it allows relating viscosity measurements to heat capacity and configurational entropy data (Urbain, 1972; Wong and Angell, 1976; Scherer, 1984, Richet, 1984). The Adam-Gibbs theory and equation will be discussed in more details in the next section. No other equation allows relating thermodynamic to dynamic variables in a simple way.

The variations observed and described previously for sodium silicate glasses are similar for all alkali oxides, as for instance visible for K_2O in Figure 6B. Generally, the addition of a few percent of alkali in silica glass produces an important viscosity decrease of a few orders of magnitude (Lecko et al., 1977). Ten mole percent of alkali oxides decreases the viscosity by ~ 10 orders of magnitude at 1400 K, this decrease being larger at low temperature near the glass transition temperature. After 10 mol% of added alkali oxides, the decrease in viscosity with increasing alkali content becomes less pronounced and the viscosity varies almost linearly at high temperature. These viscosity variations for more than 10% of added alkali are well correlated with increases in thermal expansion and heat capacity of the liquid which vary almost linearly as a function of chemical change at high temperature (Bockris et al., 1955; Lange and Carmichael, 1987; Lange and Navrotsky 1992).

2.2 ideal mixing: mixing alkali or alkaline-earth in silicate glasses and melts

In section 2.1, we observed that molar volumes, as well as viscosity and glass transition temperature of alkali and alkaline-earth silicate glasses all changes with the addition of network modifiers. Mixing different metal cations (like Ca and Mg, or Na and K) in silicate glasses results in large variations in their properties, such as large decreases of their glass transition temperatures or large increases in their electrical conductivity (Day, 1976). This is the so-called mixed alkali effect, MAE, reviewed many times (Richet, 1984, Neuville and Richet, 1991, Allward and Stebbins, 2004; Cormier and Cuello, 2013, Le Losq and Neuville, 2013, Le Losq et al., 2017, Bødker et al., 2020) in the literature for alkali silicate glasses since the work of Day (1976).

In this section, we show and introduce some terminology about mixing elements and their effect on the structure and macroscopic properties of melts and glasses. Ca and Mg are very important elements in Earth and material sciences, and it thus it is particularly important to understand how they mix in silicate melts. To understand this, we can look at how viscosity varies along the CaSiO_3 - MgSiO_3 binary. Neuville and Richet (1991) have shown that at constant temperature, the viscosity of $(\text{Ca,Mg})\text{SiO}_3$ composition is always lower than the viscosity of the end-member, CaSiO_3 or MgSiO_3 (Figure 7A).

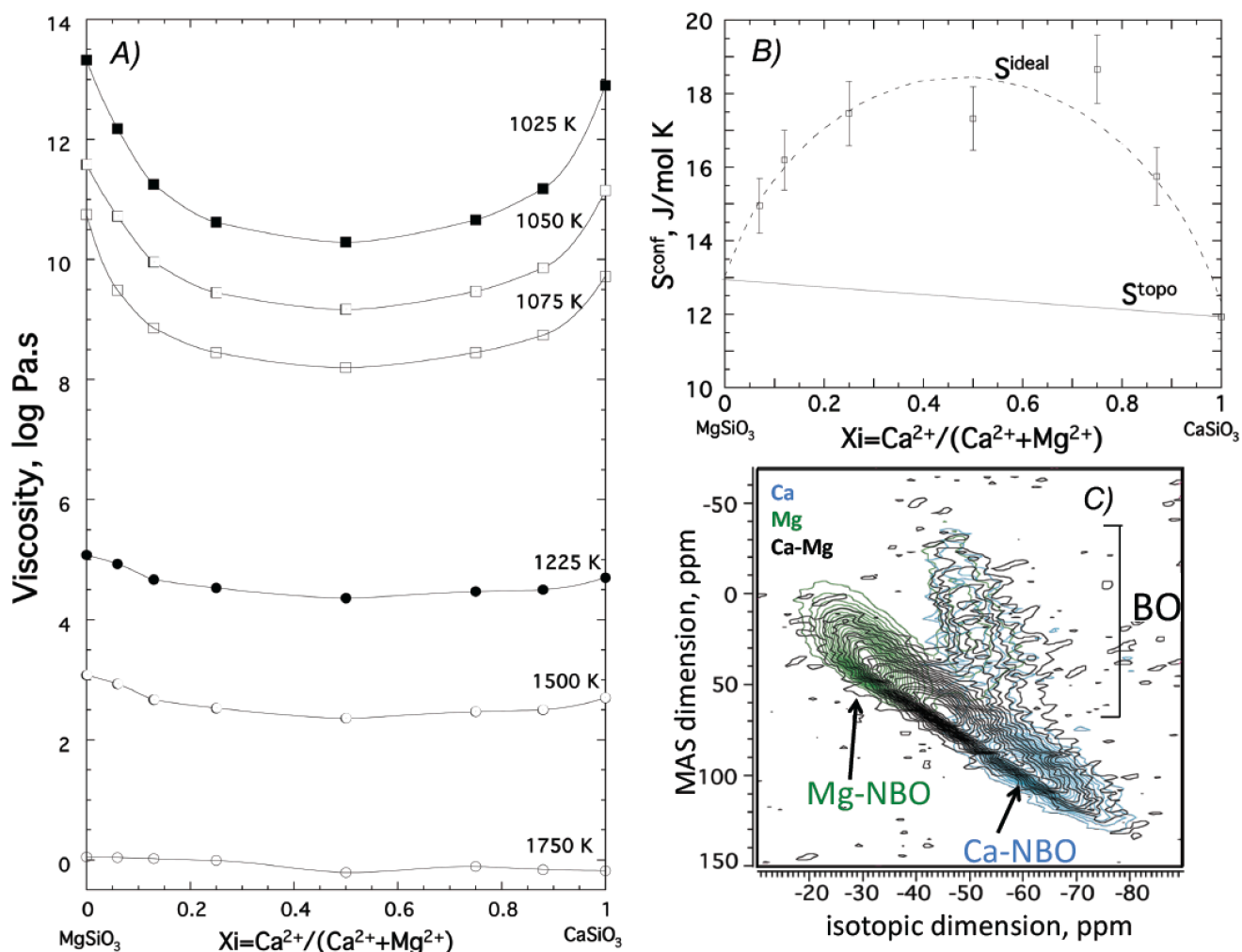


Figure 7: A) Viscosity versus $x_i = \text{Ca}^{2+}/(\text{Ca}^{2+} + \text{Mg}^{2+})$ at constant temperature and B) configurational entropy for Ca/Mg mixing data from dashed line correspond to the total configurational entropy and full line to the topological variation of the entropy see text (Neuvill and Richet, 1991) C) ^{17}O triple-quantum NMR spectrum redrafted from Allwardt and Stebbins (2004).

By using the Adam-Gibbs equation, it is possible to determine the configurational entropy, S^{conf} . Adam and Gibbs (1965) theory of relaxation processes is based on the idea that transporting matter in a viscous liquid requires a cooperative change in the fluid configuration. A liquid with zero configuration entropy would be analogous to a perfect crystal, and no material displacement could occur since a single configuration would be available. The viscosity would then be infinite. If only two configurations were possible, a movement of matter could only occur if all the atoms of the liquid changed position simultaneously. The probability of such a cooperative movement would be very low and viscosity extremely high, but no longer infinite. More generally, as S^{conf} grows, cooperative configurational changes can occur independently of each other in increasingly smaller volumes of the liquid. At the same time, viscosity decreases and is predicted by:

$$\log \eta = A_e - B_e/TS^{conf}(T), \quad (6)$$

where A_e is a pre-exponential term and B_e a measure of the Gibbs-free energy barriers hindering configurational rearrangements in the liquid. $S^{conf}(T)$ can be written as:

$$S^{conf}(T) = S^{conf}(T_g) + \int_{T_g}^T \frac{C_p^{conf}(T)}{T} dt, \quad (7)$$

And

$$C_p^{conf}(T) = C_{pl}(T) - C_{pg}(T_g), \quad (8)$$

where C_{pl} and C_{pg} are the heat capacity of the liquid and the glass.

C_p^{conf} is the melt configurational heat capacity, equal to the difference between the heat capacity of the liquid C_{pl} and the heat capacity of the glass at T_g , $C_{pg}(T_g)$ (Richet et al., 1986). Those values can be measured or modelled with the existing parametric equations (Stebbins et al., 1984; Richet and Bottinga 1984; Richet 1987; Tangeman and Lange 1998; Russell and Giordano 2017), see for examples and details Neuville (2005, 2006), Le Losq et al. (2014) and Le Losq and Neuville (2017). One may also note that $C_{pg}(T_g) \sim 3R$, the Dulong and Petit limit (Petit and Dulong, 1819), such that it can be easily estimated from this simple calculation (Richet, 1987). The configurational heat capacity is a very important term, in particular, the higher C_p^{conf} is, the faster S^{conf} increases with T, and the lower the B_e/S^{conf} ratio.

The importance of these differences between configurational C_p is illustrated in Figure 1, where in the case of $\text{NaAlSi}_3\text{O}_8$, the liquid C_p^{conf} represents less than 10% of that of the glass, and deviations from Arrhenius behavior are small. On the contrary, CaSiO_3 or $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ have a C_p^{conf} representing 30% of the C_p of the glass, and the viscosity versus $1/T$ curves depart significantly from a straight line. Because of the importance of C_p^{conf} , it is therefore imperative to know the heat capacity of liquids and glasses for a given temperature and chemical composition. We can note that the configurational heat capacity $C_p^{conf}(T)$ depends essentially on the heat capacity of the liquid, C_{pl} , and in the case of alkaline-earth silicate melts, C_{pl} is independent of T and varies mostly linearly with melt composition (Stebbins et al., 1984, Richet and Bottinga, 1985, Lange and Navrotsky, 1992). This explains the pseudo-linear variations of viscosity observed at high temperature (1750 K, Figure 7A).

However, near T_g , the melt viscosity depends strongly on $S^{conf}(T_g)$, which shows large non-linear variations with melt composition. This explains the large, non-linear viscosity changes at supercooled temperature upon mixing alkaline-earth elements in silicate melts (Figure 7A). Using equation 6 in conjunction with viscosity data and heat capacity values (from data or models), S^{conf} can be calculated. S^{conf} along the CaSiO_3 - MgSiO_3 binary is plotted in Figure 7B. To reproduce S^{conf} variations with composition, it is necessary to look at the contributions to $S^{conf}(T_g)$ in equation 7. S^{conf} records topological contributions from the melt structure (bond angle and inter-atomic distance distributions, etc.) as well as

excess entropy arising from chemical mixing effects. It thus is common to express S^{conf} as the addition of those two sources (Neuville and Richet, 1991):

$$S^{conf}(T_g) = S^{topo} + S^{mix} \quad (9)$$

with S^{mix} the term embedding all mixing contributions and S^{topo} the entropy arising from the topology of the glass network. The topology of the network for the two end-members is distinctly different and is also different compared to the crystalline analogues.

S^{topo} , the topological configuration entropy can be approximated as a linear variation of the configurational entropy of the end-members and calculated as $\sum x_i S^{conf}_i(T_g)$, with $S^{conf}_i(T_g)$ the configurational entropy of CaSiO_3 and MgSiO_3 glasses (Neuville and Richet, 1991). This S^{topo} term corresponds to a mechanical mixing between the two end-members without necessary chemical or physical interaction. S^{topo} can be expressed from the glass structure and varies linearly with composition (Le Losq and Neuville, 2017).

S^{mix} corresponds to a chemical mixing term that can be ideal or non-ideal. In the case of Ca/Mg mixing, the simplest hypothesis that can be made is that $S^{mix} = S^{ideal}$, and thus can be written as:

$$S^{ideal} = -n \sum R x_i \ln x_i, \quad (10)$$

where R is the perfect gas constant and n is the number of atoms exchanged per formula units, 1 in this case, $x_i = \text{Ca}^{2+}/(\text{Ca}^{2+} + \text{Mg}^{2+})$. From a thermodynamic point of view, the ideal solution implies that end-members components mix randomly in solution.

The excellent agreement (Figure 7B) between entropies obtained from viscosity measurements and values predicted using eqs 6 to 10 shows that the ideal mixing hypothesis for Ca/Mg is consistent with the data. It should be noted that the mixing effect is predominant near T_g . The configuration entropy of the intermediate compositions is therefore much higher than that of the end-members, explaining the minimum in viscosity observed upon Ca/Mg mixing. When temperature increases, the entropy of the end-members increases rapidly and S^{mix} , which is constant, eventually becomes small compared to $\sum x_i S^{conf}_i(T_g)$ in eq. (9). Therefore, at high temperatures, S^{mix} is negligible and this results in a quasi-linear variations in viscosity, driven by the linear variation of C_{pl} upon Ca/Mg mixing.

These macroscopic conclusions about Ca/Mg mixing are in very good agreement with nanoscale investigations made by ^{17}O NMR (Allwardt and Stebbins, 2004) and neutron diffraction and RMC (Cormier and Cuello, 2013). From neutron diffraction, the silicate network shows small but significant changes with the Mg/Ca exchange and a significant intermixing of Ca and Mg can be observed (Cormier and Cuello, 2013). From another point of view, ^{17}O NMR spectra of the diopside glass $(\text{CaMg})\text{SiO}_3$ shows that only one broad non-bridging oxygen atom (NBO) peak is visible, and encompasses the entire range of chemical shifts ranging from Ca-NBO to Mg-NBO. Comparison of the isotropic projections from 3QMAS NMR to 1D spectra predicted using a random model show that Ca/Mg mixing in Ca/Mg-

silicate glasses is generally highly disordered (Allwardt and Stebbins, 2004), except maybe near the glass transition temperature where a more ordered glass is possible as shown by neutron diffraction made on CaSiO_3 glass by Gaskell et al. (1985).

Validity of the viscosimetry approach to determine $S^{\text{conf}}(T_g)$

We have just shown that it is possible to determine configurational entropies using viscosity measurements and the AG theory. Are those values of $S^{\text{conf}}(T_g)$ truly representative of the glass residual entropy? To answer this question, we can compare $S^{\text{conf}}(T_g)$ obtained from calorimetric measurements to those obtained from fitting viscosity data. Data from 11 very different compositions, including B_2O_3 , SiO_2 or KAlSi_3O_8 , show very good agreement between $S^{\text{conf}}(T_g)$ values retrieved from viscosity measurements or calorimetric measurements (Figure 8). This thus validates the use of eq. 6 to determine $S^{\text{conf}}(T_g)$ of melts.

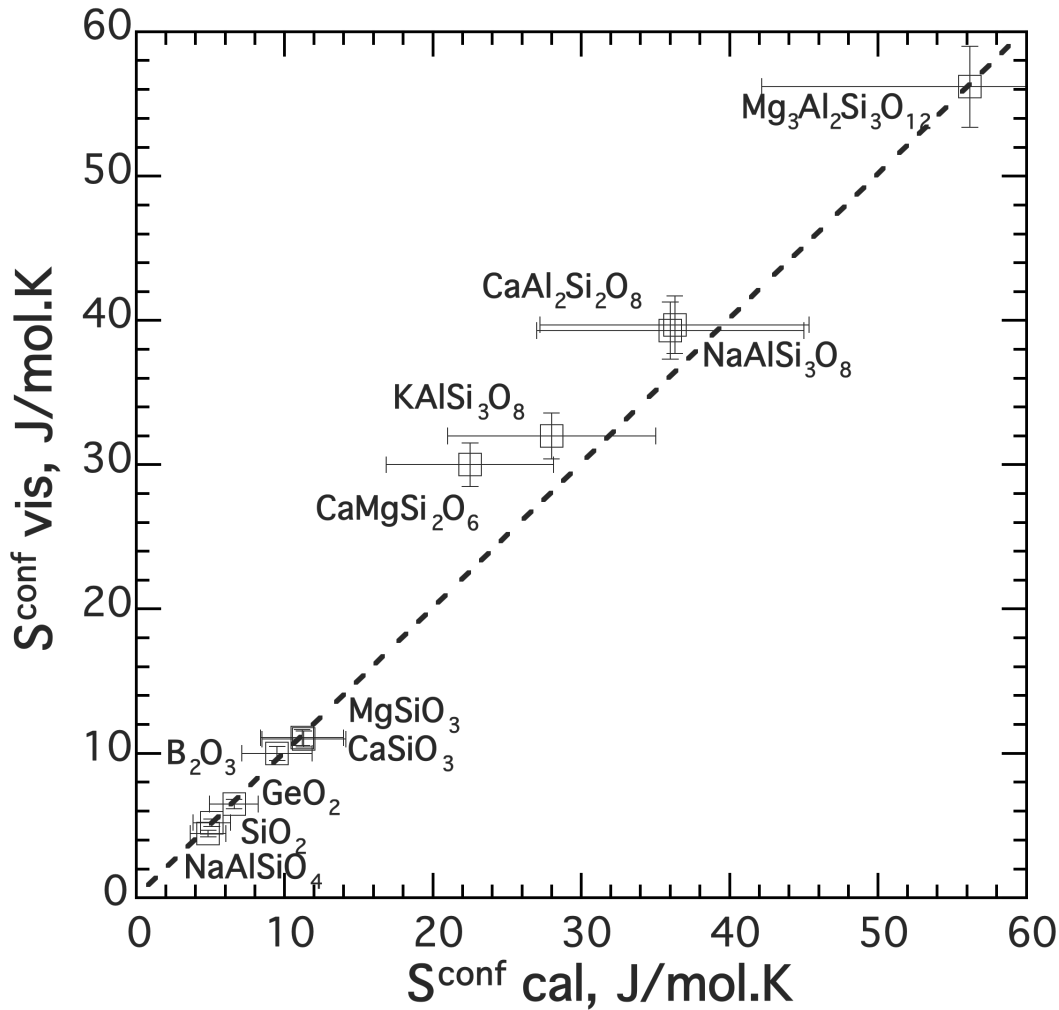


Figure 8: viscosimetry configurational entropy versus calorimetry configurational entropy for 10 different oxide glasses for which determination of calorimetric measurements are possible. Calorimetric

data are from Richet (1984) for SiO_2 , $\text{NaAlSi}_3\text{O}_8$, KAlSi_3O_8 , de Ligny and Westrum (1996) for GeO_2 and B_2O_3 , Richet et al. (1991) for NaAlSiO_4 , Tequi et al. (1991) for $\text{Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$, and Richet et al. (1986) for MgSiO_3 , CaSiO_3 , $\text{CaMgSi}_2\text{O}_6$ and $\text{Ca}_2\text{Al}_2\text{Si}_2\text{O}_8$. Viscosity data are from Hetherington et al. (1964) and Urbain et al. (1982) for SiO_2 , Neuville and Richet (1991) and Le Losq and Neuville (2013) and Le Losq et al. (2017) for alkali and alkaline-earth silicate and aluminosilicate compositions, from Fontana and Plummer (1966) for GeO_2 , and from Eppler (1966), Macedo et Litovitz. (1965), Napolitano et al. (1965) for B_2O_3 .

In the case of Na/K mixing in silicate melts, only one set of measurements is available from Poole (1948). Using those data, Richet (1984) proposed that an ideal, random Na/K mixing occurs in silicate melts, and viscosity can be modelled using the ideal mixing calculation of S^{conf} ; some differences between calculated and measured viscosity in the supercooled silicate melts are visible in Richet (1984) work, however they arise from the way the data and model are represented. The more recent study of Le Losq and Neuville (2017) resolved this, showing that it is clearly possible to model the viscosity of Na-K silicate melts while assuming a random mixing of the alkalis (see section 2.4).

2.3 Mixing alkali and alkaline-earth elements in silicate glasses

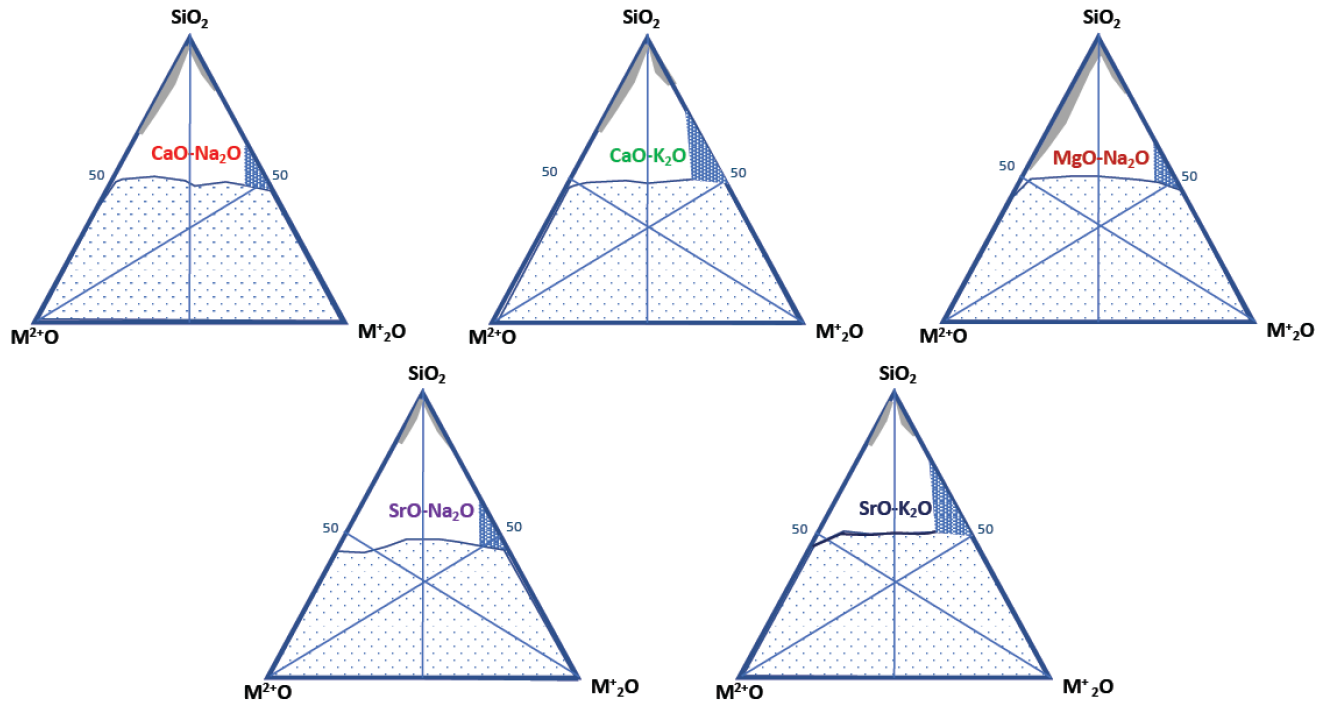
Most industrial and geologic glasses are silicate and aluminosilicate glasses containing a mixture of alkali and/or alkaline-earth elements. Igneous rocks (Morey and Bowen, 1925) as well as nearly 90 % of industrial glasses contain mixtures of alkali and alkaline-earth elements. We can cite the particular case of soda-lime silica glasses, extensively used for the construction of conventional windows and container glasses. The importance of $\text{M}^{2+}\text{O}-\text{M}^+\text{O}-\text{SiO}_2$ glasses (with M^+ an alkali and M^{2+} an alkaline-earth element) further increased recently because they also are used for the development of bioactive glasses (Hench, 1991 and Kim et al. 1995, Clupper and Hench, 2003, Hill and Brauer, 2011; Brauer, 2015). Given the importance of alkali and alkaline-earth metal cations in glasses, many studies were performed to understand the $\text{M}^{2+}\text{O}-\text{M}^+\text{O}$ mixing in silicate glasses. They are essentially focused on the glass mechanical properties or devitrification (Morey, 1930; Frischat and Sebastian, 1985; Koike and Tomozawa, 2007, Koike et al., 2007), on glass structure (Buckerman and Müller-Warmuth, 1992, Jones et al., 2001; Lockyer et al., 1995; Lee and Stebbins, 2003; Lee et al., 2003, Neuville 2005, 2006, Hill and Brauer, 2011; Brauer, 2015; Moulton and Henderson, 2021 and reference therein) or thermodynamic properties (Natrup et al., 2005; Neuville 2005, 2006, Richet et al., 2009b; Inaba et al., 2010, Sugawara et al., 2013). Given the above-mentioned importance of $\text{M}^{2+}\text{O}-\text{M}^+\text{O}-\text{SiO}_2$ glasses, we will in this section bring new insights to the effect of $\text{M}^{2+}\text{O}-\text{M}^+\text{O}$ mixture on the properties of silicate glass and melt.

2.3.1 Glass formation domains

The substitution of alkali elements by alkaline-earth elements allows extending the glass forming domain of silica glasses to lower silica contents (Figure 9), and, more generally, it allows improving many glass properties (Moore and Carey, 1951, Imaoka and Yamazaki 1963, Tomozawa, 1978; 1999; Gahay and Tomozawa 1989, 1984, Neuville, 2005, 2006). For $M^{2+}O$ and M^+O_2 bearing silicate glasses, the vitrification domain starts near 50 mol% silica in most cases. For Mg silicates, it starts at 45.7 mol% (Moore and Carey, 1951). It was generally accepted that the presence of Mg-O bonds allows preserving somehow the continuity of the silicate network, with Mg acting to some extent as a network former (Moore and Carey, 1951), an idea in agreement with the low coordination number of Mg discussed previously. In the case of CaO-Na₂O-SiO₂ glasses, the first phase diagram was made by Morey and Bowen (1925). It revealed the domain where glasses can be formed, but also an immiscibility domain in which phase separation is observed. This immiscibility domain is very large near the CaO-SiO₂ binary, extending between 62 and 100 mol% of silica. It decreases rapidly with addition of Na₂O in the glass (Morey and Bowen, 1925, Greig 1927, Hudon and Baker, 2002a,b). In the immiscibility domain, different phases exist: cristobalite, SiO₂, wollastonite CaSiO₃, sodium disilicate Na₂Si₂O₅, devitrite Na₂Ca₃Si₆O₁₆, and a non-defined compound Na₂Ca₂Si₃O₉. The proportion of these phases vary as a function of the bulk composition and the cooling rate (Morey and Bowen, 1925, Greig 1927, Tomozawa, 1978; 1999; Gahay and Tomozawa 1989; Hudon and Baker, 2002a,b; see also Schuller chapter in Neuville et al., 2017).

Similar observations can be made in the other M^{+}_2O - $M^{2+}O$ -SiO₂ systems, with $M^+=Li^+$, Na^+ , K^+ and $M^{2+}=Mg^{2+}$, Ca^{2+} , Sr^{2+} , Ba^{2+} . Generally, a pyrosilicate phase, $M^{2+}SiO_3$, and one or two phases $M^{+}_yM^{2+}_xSi_zO_u$ crystallize at high silica contents (> 70 mol%, x,y,z,u are different proportions of each crystalline phases), except for $M^{2+}=Ba^{2+}$ where the glass forming domain is large but mostly unconstrained. Bender et al. (2002) achieved the synthesis of Ba silicate glasses up to the BS7 (BaO/SiO₂=7 in mol %) composition, and Frantz and Mysen (1995) investigated some of those glass compositions by Raman spectroscopy.

To summarize, in all of the M^{+}_2O - $M^{2+}O$ -SiO₂ ternary diagrams, the glass forming domains are relatively small and essentially exist at silica contents higher than 50 mol%, with exception of barium silicates. In the case of sodium and potassium silicates, a large zone of high hygroscopicity exists below the glass forming domain, typically between 66 and 50 mol % of silica for sodium silicate glasses and between 70 and 50 mol % for potassium silicate glasses (Tomozawa, 1978; 1999; Gahay and Tomozawa 1989). In general, one should be careful with potassium silicate glasses that are always highly hygroscopic.



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Figure 9: The upper part of the ternary diagram corresponds to the glass forming region of glass between $M^{2+}O$ and M^+_2O , with $M^{2+}=Mg^{2+}$, Ca^{2+} , Sr^{2+} and $M^+=Na^+$, K^+ (redrafted from Imaoka and Yamazaki, 1963 and Levin et al., 1964). The white areas are the glass making domain, the light dotted areas are the crystallization domains, the dark dotted lines correspond to the glass making domain, but strongly hygroscopic, and the gray areas to unmixed glass.

2.3.2 Viscosity variations

Figure 10A shows viscosity measurements of silicate melts for M^{2+} -Na mixing with $M^{2+} = Ca^{2+}$, Sr^{2+} . Ca- and Sr-silicate melts show similar behavior. The addition of those elements produces a large increase in melt viscosity at constant temperature, in agreement with previous measurements (English, 1923). The lack of measurement between 10^2 and 10^9 Pa·s results from the very rapid rate of crystallization in this viscosity-temperature range for mixed alkali and alkaline-earth silicate glasses (Meiling and Uhlmann, 1977; Mastelaro et al. 2000, Neuville, 2005, 2006). In the Figure 10B, we clearly observe that addition of 10 mole % Na_2O in an alkaline-earth silicate melt has a stronger effect on melt viscosity than addition of 10 mole % $M^{2+}O$ to a soda-silicate melt composition. This indicates that a sodium silicate glass network can incorporate alkaline-earth elements more easily than the opposite. A striking difference exists in the behavior of the viscosity between pure Na-silicate melts and pure alkaline-earth silicate melts (CN60.00 or SN60.00). Indeed, at constant viscosity and near T_g (Figure 10B), i.e. near 1000 K, the viscosity of the end-member compounds differ by 10 orders of magnitude. This difference decreases

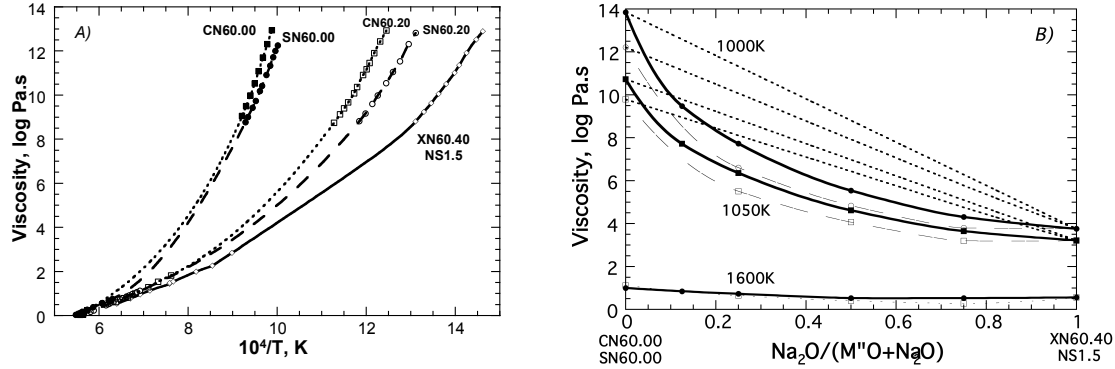


Figure 10: A) Viscosity versus $1/T$ for silicate melts ranging from the pure alkaline-earth silicate melt) to pure soda silicate (XN60.40) composition names are: XN60.Y, X=C or S for the Ca or Sr system, Y correspond to Na_2O content, 60 to SiO_2 in mole percent, and $X = 100 - (60 + Y)$ all in mole percent. B) Viscosities of Na/M^{2+} silicate melts versus $\text{Na}_2\text{O}/(\text{Na}_2\text{O} + \text{M}^{2+}\text{O})$ at constant temperature, $\text{M}^{2+} = \text{Ca}, \text{Sr}$, dashed lines correspond to the linear viscosity variation at 1000 K, 1050 K and 1600 K, solid and open symbols correspond to the $\text{CaO}-\text{Na}_2\text{O}-\text{SiO}_2$ and $\text{SrO}-\text{Na}_2\text{O}-\text{SiO}_2$ systems, respectively compiled from Neuville (2006, 2005).

The viscosity variations observed in Figure 10 can be understood using equations 6 to 9. It is possible to determine how configurational entropy varies upon the mixing of Na_2O with M^{2+} in the glasses (Figure 11A). S^{conf} can be decomposed in S^{topo} and S^{mix} (eq. 9). S^{topo} is equal to the sum of endmembers. S^{mix} , in the case of M^+/M^{2+} mixing, has a complex expression that depends on the way the cations mix into the glass structure.

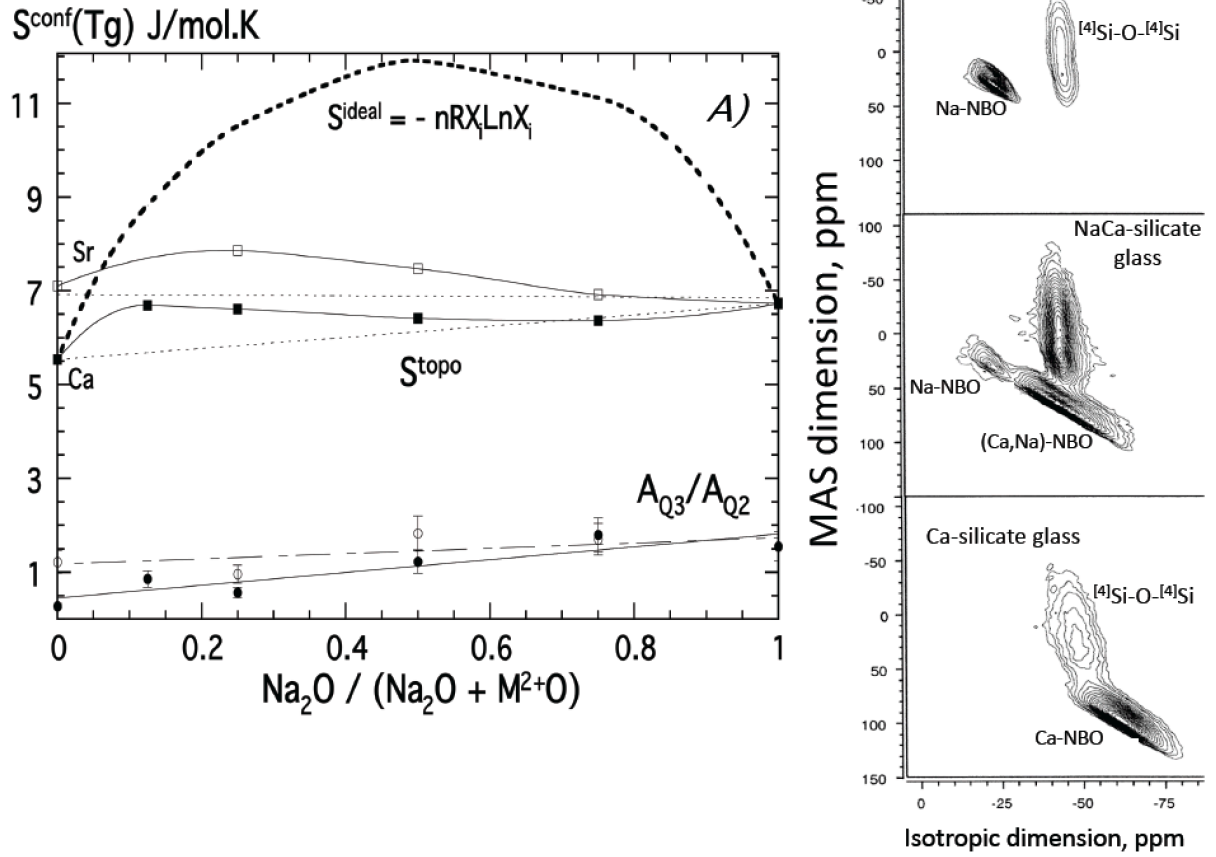


Figure 11: A) $\text{Na}_2\text{O}/(\text{Na}_2\text{O} + \text{M}^{2+}\text{O})$ dependence of configurational entropy, $S^{\text{conf}}(T_g)$, A_{Q3}/A_{Q2} ratio for in black symbol CaO and in open symbol for SrO (entropy are in J/mol.K and A_{Q3}/A_{Q2} without units and are calculated from the Q^3 and Q^2 areas of the Raman spectra, see Neuville (2005, 2006) for more detail). B) ^{17}O 3QMAS MAS spectra for Ca-Na silicate glasses at 9.4 T redrafted from Lee and Stebbins, (2003).

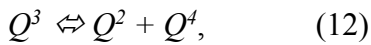
Between the calcium and sodium silicate glasses, $S^{\text{conf}}(T_g)$ varies non-linearly with $\text{Na}_2\text{O}/(\text{Na}_2\text{O} + \text{CaO})$. It increases rapidly with increasing $\text{Na}_2\text{O}/(\text{Na}_2\text{O} + \text{CaO})$ to 0.2, then it slightly decreases with further substitution Ca^{2+} by Na^+ . Similar observations are made in the $\text{SrO}-\text{Na}_2\text{O}-\text{SiO}_2$ system (Neuville, 2005). The observed variations strongly depart from the predictions made by the ideal model given in equation 10 (Figure 11A). In the case of Na^+ and Ca^{2+} , or of Na^+ and Sr^{2+} , it is clear that, as their $S^{\text{conf}}(T_g)$ variations are close to linear trends, those atoms do not mix randomly in the silicate melts. The fits of $S^{\text{conf}}(T_g)$ presented in Figure 11A thus require the use of a non-ideal mixing model. To model S^{mix} in the case of a non-ideal mixing of alkali and alkaline-earth elements, Neuville (2005, 2006) proposed to use the equation:

$$S^{\text{mix}} = W_{NM} x_n y_c^3, \quad (11)$$

where $x_n = \text{Na}_2\text{O}/(\text{Na}_2\text{O} + \text{M}^{2+}\text{O})$ and $y_c = 1 - x_n$, and W_{NM} is a constant.

2.3.3 Structure of mixed glasses

Equation (11) implies a significant change in the residual entropy that originates from a re-arrangement of the glass structure as Ca^{2+} is replaced by Na^+ . Spectroscopic methods that probe the glass structure allow us to investigate such structural changes. ^{23}Na and ^{17}O Nuclear Magnetic Resonance (NMR) spectroscopy data shown a non-random distribution in the case of Na^+ - Ca^{2+} mixing, with in particular a prevalence of Na^+ - Ca^{2+} pairs in the system (Figure 11B), and several partially resolved non-bridging oxygen peaks such as Na^+ -NBO, and mixed peaks (Na^+ , Ca^{2+})-NBO (Lee and Stebbins, 2003). The observed fractions of Na^+ -NBO are smaller than those predicted by random distributions of Na^+ and Ca^{2+} , suggesting preference for dissimilar pairs (Lee and Stebbins, 2003) which is also in good agreement (Florian et al., 1996; Lee 2005, Lee and Stebbins 2002 and Lee et al., 2003). In the previous chapter, Raman spectroscopy soda and lime silicate glasses were shown in Figure 9. Different Q^n contributions are visible in those spectra. From these Q^n contributions, the ratio of the area between Q^3 and Q^2 units, A_{Q3}/A_{Q2} can be calculated. Results of such calculation are plotted in Figure 11A. It shows that the fraction of Q^3 units in the glass increases with increasing the $\text{Na}_2\text{O}/(\text{Na}_2\text{O}+\text{M}^{2+}\text{O})$ ratio. M^{2+}O silicate glasses are thus richer in Q^2 and Q^4 units than alkali silicate glasses (Frantz and Mysen, 1995, Neuville 2005, 2006), because increasing the ionic field strength of metal cation in the glass shifts to the right the equilibrium of the reaction:



This observation will be used in section 2.4 for highlighting how one can build models of the structure, entropy and viscosity of silicate melts. Huang et al. (2016) investigated Ca_3SiO_5 composition by using $^{17}\text{O} \rightarrow ^{29}\text{Si}$ CP-HETCOR NMR experiment and they proposed that free oxygen can also play a small role structure and properties in the mixed glasses, what is possible but this experiment must be considered with caution, firstly it is a particularly difficult composition to manufacture and secondly, only the nominal compositions are given and not those analyzed.

2.3.4 Network polymerization

The substitution of 30 mol% of CaO by Na_2O in an alkaline-earth silicate glass results in a moderate increase of $\sim 0.9 \text{ J mol}^{-1} \text{ K}^{-1}$ of $S^{\text{conf}}(T_g)$ (Figure 11A). This indicates an increase in the structural disorder in the glass, possibly linked to change in the glass network polymerization. Structural changes that may account for the entropy behaviour may be discerned in the Raman spectra of the glasses. In particular, the ratio of the areas of Raman bands assigned to Q^3 and Q^2 units, A_{Q3}/A_{Q2} , varies with $\text{M}^{2+}\text{O}/(\text{M}^{2+}\text{O}+\text{M}^{2+}\text{O})$, indicating a change in the Q^n speciation distribution (Neuville, 2006). There is a distinct increase of this ratio with adding Na_2O in the alkaline-earth silicate glasses. In the case of Na/Ca

and Na/Sr mixing, Raman spectroscopy observations shown in Figure 11A are in good agreement with previous observations by Raman spectroscopy (McMillan, 1984; McMillan et al., 1982; Frantz and Mysen, 1995) and by ^{17}O NMR (Maekawa et al., 1991; Buckerman and Müller-Warmuth, 1992; Jones et al., 2001, Lee and Stebbins, 2003, Lee et al., 2003), all showing a decrease in Q^3 fraction upon substitution of alkali by alkaline-earth element in silicate glasses. Moreover, the increase in Q^2 units with adding alkaline-earth elements in soda silicate glasses can be correlated to increasing Boson peak frequency and intensity (Neuville, 2006).

Note that below 250 cm^{-1} , there is only a scattering continuum and the Raleigh tail of the exciting line, except at very low frequency where there is the so-called boson peak (Malinovsky and Sokolov, 1986; Buchenau et al., 1986). This peak has been ascribed to excitations associated with rotational motions of almost rigid tetrahedra (Buchenau et al., 1986; Helhen et al., 2000, 2002). Helhen et al. (2002) consider that this peak, observed for different silicate glasses, increases in intensity and shifts to higher frequency with higher distortion of the SiO_4 tetrahedra. Several others interpretations are proposed: - Champagnon and collaborators proposed that Boson peak can be link with structural homogeneity in the glass (Champagnon et al., 2009), and the other hand Chumakov and collaborators proposed that Boson peak can be attributed to some fluctuation of disorder in glass (Chumakov et al., 2014) and finally Schirmacher and collaborators proposed an intermediate interpretation by considering that the Boson peak can corresponds to the spatial fluctuations of elastic constants caused by the structural disorder of the amorphous materials (Schirmacher and Ruocco, 2020, Schirmacher et al., 2015). If we consider that molecular dynamic simulations provided by Cormack and coworkers suggest that Si-NBO bond length is about 0.05Å shorter than the Si-BO bond length (Henderson, 1995; Cormack and Du, 2001; Du and Cormack 2004)⁵, we can propose that Q^2 units are more distorted than Q^3 or Q^4 units. This implies that Q^4 and Q^0 units are regular tetrahedra with Q^4 bigger than Q^0 and they are less distorted than the Q^1 , Q^2 , Q^3 units species (Nakamura et al., 2013). It is probable that Q^2 units are the more distorted tetrahedra with 2 distances T-BO and 2 distances T-NBO which can be in good agreement with the variation of the frequency of boson peak. So, we proposed that the variation on the Boson peak observed as a function of chemical change in simple systems with cation substitution can be attributed to tetrahedra distortion as observed on SiO_2 glass with pressure. This variation can also result of structural heterogeneity produce

⁵ Du and Cormack (2004) proposed from MD simulation that when the alkali content vary between 10 and 50 mole percent the $d(\text{Si-BO})$ vary between 1.625 and 1.629Å and $d(\text{Si-NBO})$ varies between 1.559 and 1.579Å . Henderson (1995) proposed from EXAFS measurements that $d(\text{Si-O})$ equal $1.61\pm0.02\text{Å}$ and $1.58\pm0.02\text{Å}$ for NS1.5 glass which is in good agreement with MD.

by the fact that soda silicate glasses are richer in Q^3 species than lime-silicate glasses. These structural heterogeneities can also be close the idea of percolation channels proposed by Greaves (1985).

2.3.5 Fragility

Angell (1991) proposed a classification of liquids between strong and fragile: a strong liquid shows a linear variation like SiO_2 in the diagram $\log \eta$ versus T_g/T whereas this linear relation is not preserved in a fragile liquid (Figure 12). The melt fragility, m , is related to its thermodynamic properties via the relationship:

$$m = B_e / S^{conf}(T_g) * T_g * (1 + C_p^{conf}(T_g) / S^{conf}(T_g)) \quad , \quad (13)$$

In the case of the $\text{M}^+_2\text{O}-\text{M}^{2+}\text{O}-\text{SiO}_2$ system, the viscosities of the alkali- alkaline-earth silicate melts are plotted as a function of T_g/T in Figure 12. The strongest liquid is the pure soda silicate, XN60.40 (NS1.5), whereas the most fragile liquids are the pure alkaline-earth silicate melts. The Ca- and Sr-silicate melts have similar fragility, despite the Ca-silicate melt being more viscous than Sr-silicate melt. Intermediate chemical compositions mixing Na^+ with Ca^{2+} or Sr^{2+} present an intermediate behavior between end-members. The behavior of the melt fragility correlates well with melt C_p^{conf} values, which decrease of $\sim 20\%$ upon replacing M^{2+} with Na ($\text{M}^{2+} = \text{Ca}^{2+}$ or Sr^{2+} ; Neuville, 2005, 2006). This observation on the decrease of the C_p^{conf} is in good agreement with the increase of the $\Delta C_p = C_{pl} / C_{pg}(T_g)$ which varies between 1.2 for the Na-silicate glass up to 1.27 and 1.29 for Ca- and Sr-silicate glasses. Note that, Angell (1991) proposed that ΔC_p vary between 1.1 and 1.3 for a strong and fragile liquid respectively.

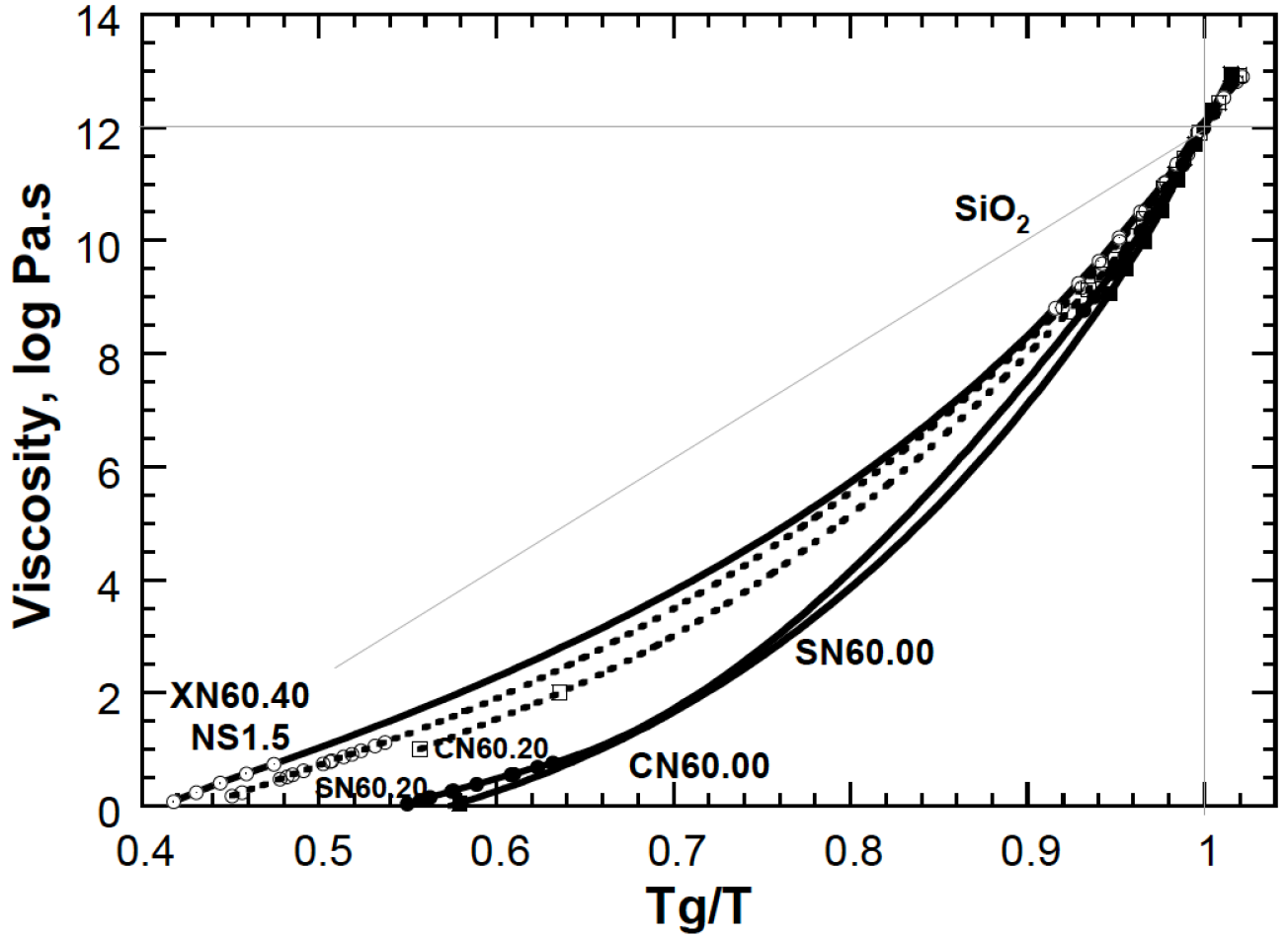


Figure 12: Viscosity of soda- M^{2+} O silicate glasses versus T/T_g . Error bars are smaller than symbol size, from Neuville (2005, 2006).

2.3.6 Melt configurational entropy, fragility, Q^n speciation, mobility and network connectivity

The distribution of cations in the glass network can have important implications for transport properties such as conductivity, viscosity and fragility. In the case of the Na/Ca substitution in silicate melts, Roling and Ingram (2000) have shown that the mobility of Ca^{2+} cations is enhanced upon their replacement by Na^+ . This replacement of Ca^{2+} by Na^+ produces a positive coupling effect on the movements of divalent cations that correlates well with the increase in $S^{conf}(T_g)$ (Figure 11A). Conversely, the activation energies related to movements of Na^+ is slightly changed. The cation mobility strongly depends on the glass structure, and the structural changes taking place during the Na/Ca substitution in the Ca-silicate glass enhance the mobility of Ca^{2+} . The modification in network topology will provide new types of empty sites that could be used for Ca movements (Cornier and Neuville, 2004). The mobility of the divalent cation is then assisted by positive coupling with the more mobile monovalent cation (Magnien et al., 2004, 2008).

To summarize the observations made in the case of the mixing between alkali, M^+ , and alkaline-earth, M^{2+} , elements in silicate glasses and melts:

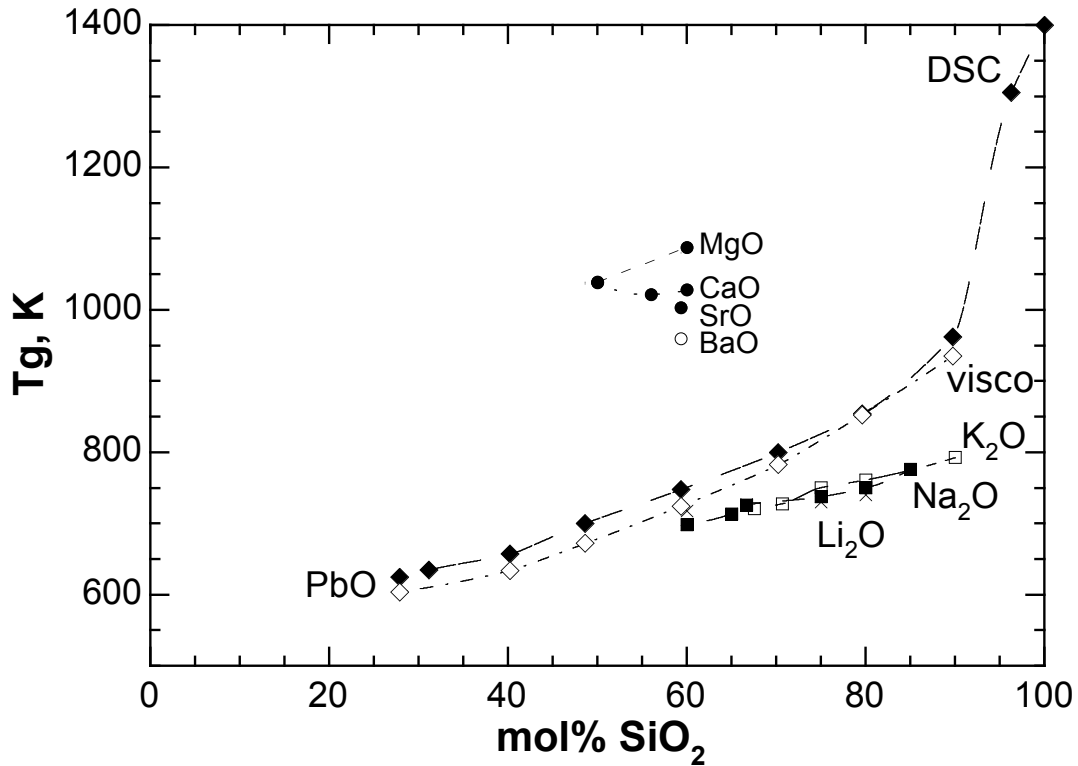
- alkali and alkaline-earth cations do not mix randomly,
- increased alkali content increases the formation of Q^3 species,
- diffusivity increase with alkaline-earth element (Roling and Ingram, 2000),
- fragility increases with alkaline-earth content,
- the distortion of the tetrahedra suggested by the boson peak variation and the distortion of the network increases with increasing MO content,
- the proportion of Q^2 species increases with $M^{2+}O$ content and in the same way, the intensity and the frequency position of the boson peak increase. The fact, that the alkaline-earth silicate glasses around than 60% of silica correspond to a mixing of Q^2 and a small amount of Q^4 species, may be the origin of the unmixing observed for more than 60% of SiO_2 in the $M^{2+}O-SiO_2$ binary system (see Figure 9). These two interconnected Q^n species can also be an image of the percolation channels (Greaves et al., 1981, Greaves, 1985, Frischat et al., 2001, Le Losq et al., 2017).

2.4 Silicate glasses and others network formers

2.4.1 Lead

The replacement of SiO_2 by PbO produces a small decrease in the glass molar volume until the Si/Pb ratio reaches a value of 0.5. Further PbO addition results in a molar volume increase. The PbO amorphous molar volume value can be obtained by extrapolation and is 25.0 cm^3 (Robie et al., 1979; see arrow in Figure 5) which is similar to the molar volume of Litharge, PbO with Pb^{2+} in four-fold coordination. Variations in the molar volume of lead silicate glasses correlate well with structural observations made via Raman spectroscopy (chapter 2, this volume, and Ben Kacem et al., 2017). For small amounts of lead oxide added to silica glass (<20 mole %), Pb is essentially in six-fold coordination and its introduction into silica results in the formation of non-bridging oxygens, clearly visible in the Q^n species distribution in the Raman spectra (Ben kacem et al., 2017). This is accompanied by a decrease in melt viscosity and glass transition temperature. At PbO concentrations higher than 20 mol%, a mixture of essentially two lead species, in 4- and 6-fold coordinations, exists. Kohara et al. (2010) and more recently Sampaio et al. (2018) mentioned that Pb can be distributed between 3-, 4-, 5- and 6-fold coordination but recent X-Ray diffraction shows that Pb is essentially distribute in 4- and 6-fold coordination (Drewitt et al., this issue). Pb in four-fold coordination can be considered as a weak network former (Fayon et al., 1998, 1999; Feller et al., 2010; Ben Kacem et al., 2017) as $[^{4}Pb]$ creates a network of interconnected tetragonal pyramid PbO_4 (Morikava et al., 1982). This is visible in Raman spectra, where an intense band at 140 cm^{-1} can be

846 assigned to Pb-O-Pb covalent bonds (see chapter 3 and Ben Kacem et al., 2017). This band at 140 cm⁻¹
 847 is clearly at higher frequency than that observed in alamosite, PbSiO₃, at 110 cm⁻¹ and attributed to Pb in
 848 6-fold coordination (Worrell and Henshall, 1978; Ben Kacem, 2017, Pena et al., 2017). The coexistence
 849 of network former PbO₄ and network modifier PbO₆ entities agree with the pair correlation function
 850 obtained from the X-ray diffraction data (see chapter 3, Figure 6B (Drewitt et al. 2021)), which show that
 851 the proportion of PbO₄ increases with decreasing SiO₂/PbO. The mixing of the tetragonal pyramid PbO₄
 852 with SiO₄ tetrahedra is probably mechanical in nature, i.e. the two polyhedra cohabit and are
 853 interconnected but their mixture does not result in an excess of entropy as indicated by the linear
 854 decreases of the glass transition temperature between 30 and 80% of SiO₂ (Figure 13). The fast decrease
 855 in T_g upon addition of ~10 mol% PbO probably relates to the network modifier role of PbO₆ octahedra,
 856 containing some NBOs in their first coordination shell.



857
 858 **Figure 13:** Glass transition temperature as a function of the fraction of SiO₂ in PbO-SiO₂ glasses.
 859 Data for Ca²⁺, Mg²⁺, Na⁺, K⁺, Li⁺ silicate glasses are original data from Neuville, Ca50 and Mg50
 860 (Neuville and Richet, 1991), Ca60 and NS1.5, NS2, NS4 from Neuville (2006), NS3 from Le Losq et al.
 861 (2014); Sr60 from Neuville (2005) and Pb-silicate glasses open circle are from viscosity measurements
 862 and full circle are from DSC measurements. (Ben Kacem et al., 2017). Note that all T_g are obtained
 863 using viscosity measurements to be comparable except those obtained by Ben Kacem with DSC.

864

865 The introduction of PbO thus produces the transformation of Q^n units into Q^{n-1} units as PbO₆ are
866 network modifier units. The appearance of these Q^{n-1} species is documented by changes in the high
867 frequency envelope (800 - 1200cm⁻¹) of the Raman spectra of the glasses (Worrell and Henshall, 1978;
868 Ben Kacem et al., 2017). The Raman spectra of PbO-SiO₂ glasses corresponds to the combination of two
869 distinct endmember Raman signals: that of a lead-silicate glass network with Q^{n-1} species, and that of an
870 amorphous lead network characterized by a strong peak at 140 cm⁻¹. The relative abundances of these
871 signals, and of their associated networks, vary as a function of the SiO₂/PbO ratio and this is in good
872 agreement with observation made using ¹⁷O NMR by Lee and Kim (2015). Lee and Kim (2015) also
873 observed a small amount of free oxygen as early mention by Sampaio et al. (2018), but this small amount
874 cannot explain the T_g and viscosity variations.

875

876 In this sense, lead is an example of an amphoteric⁶ element that plays two different roles, network
877 former or network modifier, depending on melt chemical composition. Another example of an element
878 presenting such behavior is Fe³⁺, which, depending on its coordination, plays different structural roles
879 (Moretti, 2005). Indeed, while Fe³⁺ in 4- or 5-fold coordination can be considered as network formers,
880 Fe³⁺ in 6-fold coordination behaves as a network modifier in silicate melts. It can be found in the literature
881 that such elements are “intermediate” and often aluminum is associated with this terminology. However,
882 the story is more complex than this, and Al³⁺ does not really present amphoteric behavior per se as highly
883 coordinated Al species are not surrounded by NBOs such that ^[4]Al is exclusively a network former
884 element. Its behavior is however complex, as we shall see in the sections below.

885

886 2.4.2 Al in silicate glasses

887 A well-known feature of silicate mineral structures is that Si cations can be replaced isomorphously
888 by aluminum cations, provided that either mono or divalent cations can ensure charge compensation of
889 Al polyhedra to maintain local neutralization of electrical charges in the structure. As such, those mono
890 or divalent cations are called charge compensators. However, it is also possible to obtain alumina-silica
891 glasses without the presence of charge compensator cations. From a research point of view, this is very
892 interesting because Al₂O₃-SiO₂ glasses are important materials for the understanding of how network
893 former elements mix. Al₂O₃-SiO₂ glasses have a very high liquidus (Sossman, 1933), and are difficult to
894 synthesize except by fast quench or sol-gel methods (e.g. see Wang et al., 2020). With those methods, it
895 is possible to obtain Al₂O₃-SiO₂ glasses in three different regions, near 15 mol%, 50 mol% or 60 mol%

⁶ An amphoteric element can react both as an acid or a base; see Moretti (2005) for details on this concept for silicate melts.

of silica. The molar volume of $x\text{Al}_2\text{O}_3-(1-x)\text{SiO}_2$ glasses calculated from density measurement (Wang et al., 2020) varies linearly between 27.3 up to 31 cm^3 for 60 % of Al_2O_3 (Figure 6).

From the data on Al_2O_3 (corundum), the molar volume of Al in six-fold coordination is 25.57 cm^3 (Robie et al., 1979). The molar volume of Al in five-fold coordination is unknown, but the extrapolation of the Al_2O_3 - SiO_2 line in Figure 5 yields a value of $\sim 32 \text{ cm}^3$ at 0 mol % of SiO_2 . Poe et al. (1992) proposed that Al_2SiO_5 glass is composed by 50% of $^{[5]}\text{Al}$, 40% of $^{[4]}\text{Al}$ and 10% of $^{[6]}\text{Al}$.

However, no information on the geometry of the $^{[5]}\text{Al}$ oxygen polyhedra is available: it could be a tetragonal pyramid, or a trigonal bi-pyramid. The present data do not allow assessment of this problem, but it is possible to make an assumption: the geometric shape of $^{[5]}\text{Al}$ will depend on the nearest charge compensator cation. For example, the smaller the quadrupolar moment, Cq , observed in ^{27}Al NMR is, the more $^{[5]}\text{Al}$ will look like a square base pyramid. The role of $^{[4]}\text{Al}$, $^{[5]}\text{Al}$ and $^{[6]}\text{Al}$ in the glass and melt structure will be discussed later in section 3.

2.5 Silicate melts: how can we use existing structural knowledge to model melt properties

With the above information from the traditional glass forming ranges, we understand that we know for silicate melts and glasses:

- the glass has a network of Q_n units, whose concentrations can be determined via NMR or Raman spectroscopy (Farnan and Stebbins, 1990; Stebbins et al., 1992; Dupree et al., 1986; Maekawa et al., 1991; Mysen, 1990; Mysen and Frantz, 1993; 1994; Ispa et al., 2010; Zhang et al., 1996; Sen et al., 2003; Davis et al., 2010; O'Shaughnessy et al., 2020; Nesbitt et al., 2018, 2021)
- glass and melt heat capacities, which can be calculated, and
- their rheological properties.

It has been a longstanding goal to link those observations via a single model that could predict the properties and structure of glasses and melts. A pioneering start was the work of Mysen (1995), who proposed to calculate the melt configuration heat capacity from partial heat capacity values $C_p^{conf} Q_n$ assigned to each Q^n unit with a fraction x_{Q^n} :

$$C_p^{conf} = \sum x_{Q^n} C_p^{conf} Q^n. \quad (14)$$

Links between melt thermodynamic parameters, rheology and structure were further established by subsequent publications (Neuville and Mysen, 1996; Neuville 2006; Mysen 1990, 1995, Le Losq et al., 2013, 2017, Le Losq et al., 2015, 2017; Russell and Giordano 2017). Recently, Le Losq and Neuville (2017) demonstrated how, using the Adam and Gibbs theory and Q^n values from NMR spectroscopy, it is possible to determine with a high precision the viscosity of Na_2O - K_2O - SiO_2 melts. In this section, we will show how such a model can be constructed, and what opportunities for extension

are envisioned. The application of such modeling for aluminosilicate melts will further be discussed in section 3.

Using a principle similar to that of equation (14), we may be able to calculate values of $S^{conf}(T_g)$ and B_e in eqs. 6 to 9 from partial values assigned to Q^n units in the glass, i.e. in the melt at the glass transition temperature. This is possible because B_e is temperature-independent (Adam and Gibbs, 1965), and $S^{conf}(T)$ can be estimated from $S^{conf}(T_g)$ and C_p^{conf} , which can be easily determined from viscosity data and modeled in silicate melts, respectively. This implies that we can use glass structural data to build a model for melts. However, we need to be able to model precisely how the distribution of Q^n units vary with glass composition. Fortunately, the distribution of Q^n units mostly depends on (i) the glass silica concentration, and (ii) the ionic field strength (IFS) of network modifier metal cations. Numerous ^{29}Si NMR spectroscopy data were acquired on silicate glasses and allow us to know precisely how Q^n units distribute depending on glass composition. They allow, in particular, determining the equilibrium constants of the Q^n dissociation reactions

$$Q^n = Q^{n-l} + Q^{n+l}, \quad (15)$$

which is at play in silicate glasses and melts. This explains the deviations of the Q^n distribution measured compared to silicate minerals, where binary combinations of Q^n units are observed (e.g. Q^4 - Q^3 , Q^3 - Q^2 , etc.). In glasses and melts, randomness implies that the Q^n distribution departs from the binary model. The existing ^{29}Si NMR and Raman data allow one to estimate how the melt composition influences such deviation. In Figure 14, we collected K_{eq} equilibrium constants for the dissociation of Q^3 units and of Q^2 units, and reported them as a function of the ionic field strength. In this Figure, most data are from NMR spectroscopy because data interpretation from ^{29}Si NMR 1D and 2D methods (Edén, 2012) is less subject to debate, despite some possible complications that highlight that the determination of Q^n unit distribution always is an arduous task (Pedone et al. 2010; Charpentier et al. 2013). Indeed, Raman data requires the use of Raman cross sections, determined by calibration against ^{29}Si NMR data, to transform the area of Raman peaks assigned to different Q^n units into concentrations of those Q^n units. This is a first barrier as Raman cross-sections are not straightforward to obtain, a second one being the fit of the spectra that can be a topic of debate (Nesbitt et al., 2018, 2021; Bancroft et al., 2018). The very few K_{eq} values from Raman data used in Figure 14 do not come from possibly controversial traditional peak fitting methods, but from the use of linear alternative least square methods to retrieve partial Raman spectra for the different Q^n units in binary silicate glasses, a method developed by Zakaznova et al. (2007) and Malfait et al. (2007, 2009). While it is very powerful for binary silicate glasses, it relies on composition-specific, constant Raman signal shapes for the different Q^n units. This hypothesis is valid in binary systems, but this method is limited because (i) partial Raman spectra for one binary system (e.g.

Na₂O-SiO₂) cannot be used for others and (ii) this starting hypothesis becomes void for ternary and more complex systems. This is because partial Raman spectra and Raman cross-sections differ depending on the network modifier present into the glass (e.g. Malfait et al., 2009; Woelffel et al. 2015).

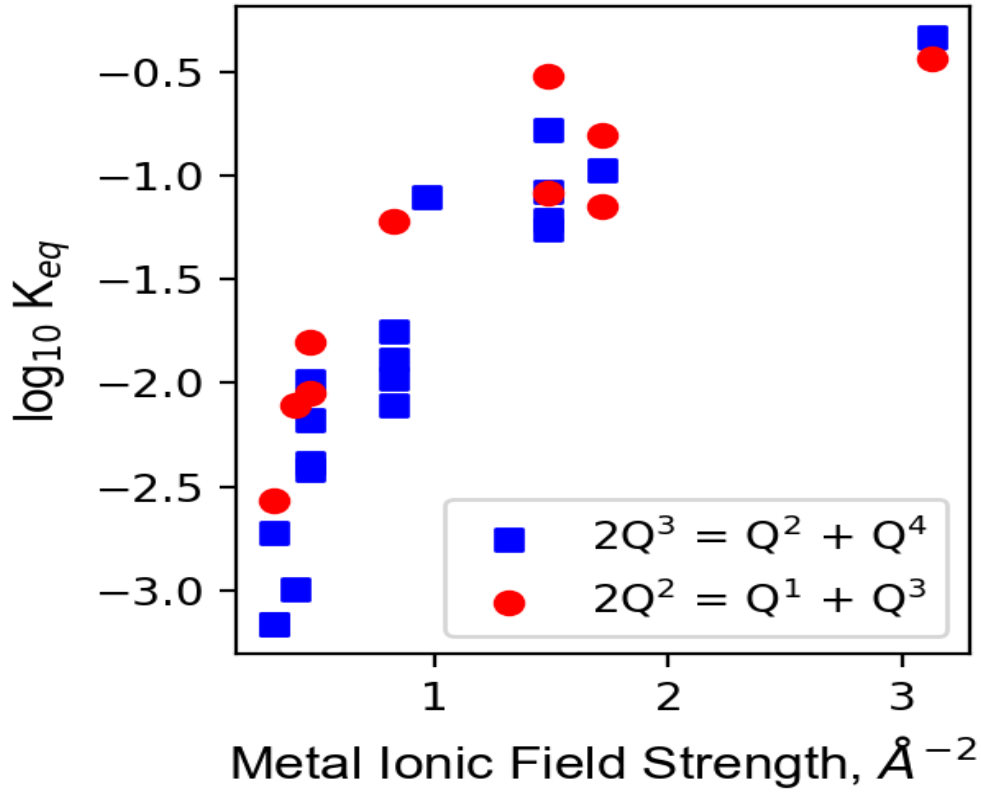


Figure 14: Equilibrium constants of the dissociation of Q^3 and Q^2 units represented as a function of the metal ionic field strength. Metal IFS was calculated assuming a coordination number of 6 and ionic radius from Whittaker and Muntus (1970). Equilibrium constants were calculated from the data of Emerson (1989), Emerson and Bray (1994), Maekawa et al. (1991), Bray et al. (1991), Zhang et al. (1996, 1997), Schneider et al., (2003), Voigt et al. (2005), Larson et al. (2006), Malfait et al. (2007), Zakaznova-Herzog et al. (2007), Sen et al. (2009), Malfait et al. (2009), Davis et al. (2010) and Herzog and Zakaznova-Herzog (2011).

An idea that arises when seeing Figure 14 is to parametrize the K_{eq} values as a function of the mean ionic field strength of metals cations in silicate glasses. Indeed, K_{eq} values for Q^3 and Q^2 units change systematically depending on the ionic field strengths of the metal cation in the glass structure. It thus seems possible to calculate K_{eq} values from the mean ionic field strengths of metal cations in the glass. For ternary and more complex silicate glasses, this relies on the fact that mixing different cations in

silicate glasses affects linearly the distribution of Q^n units. Fortunately, this is a good approximation. Indeed, from ^{29}Si and ^{23}Na MAS NMR spectroscopy, the Q^n unit distribution upon Rb-Na mixing in alkali trisilicate glasses does not systematically change (Hater et al., 1989). This was also reported during Na and Li mixing in $(\text{Na},\text{Li})_2\text{Si}_2\text{O}_5$ glasses (Ali et al., 1995, Sen et al., 1996). From ^6Li , ^7Li and ^{29}Si static NMR spectroscopy data, Bray et al. (1991) observed a linear change of the Q^3 units fraction upon mixing Li and K in disilicate glasses. The ^{29}Si static NMR data from Emerson and Bray (1994) further suggest a slight departure from linearity of the Q^3 unit fraction when mixing $\text{Na}_2\text{Si}_2\text{O}_5$ and $\text{Cs}_2\text{Si}_2\text{O}_5$ glasses. Such interpretation may be consistent with ^{17}O Dynamic Angle Spinning NMR data from mixed $\text{Na}_2\text{Si}_2\text{O}_5$ - $\text{K}_2\text{Si}_2\text{O}_5$ glasses (Florian et al., 1996), showing slightly non-linear changes of the fractions of Bridging (BO) and Non-Bridging (NBO) oxygen anions with changes in the glasses K/ (K + Na) ratios. Such observation agrees with the recent report of Le Losq and Neuville (2017), which showed that the Raman areas of Q^n units in alkali trisilicate glasses vary linearly within error bars.

From the above review, it thus should be possible to calculate K_{eq} values for Q^3 and Q^2 unit dissociation given the mean IFS of metal cations present in the silicate glass. However, despite a good correlation, scattering is visible in Figure 14. Such scattering results from second order effects that may be related to some local peculiarities of the effect of each cation on K_{eq} values as well as our lack of a perfect understanding of the coordination number of each cation. To circumvent this problem, we propose to calculate the mean K_{eq} values in silicate glasses based on the combination of K_{eq} values assigned to each cations (from their calculation in the binary $\text{M}^{x/2}\text{O}$ - SiO_2 systems):

$$K_{eq}^{glass} = \sum K_{eq}^i x_i, \quad (16)$$

where x_i is the fraction of a given cation relative to the sum of all cations. Assuming linear variations in the Q^n unit fractions upon mixing different cations, a valid hypothesis given existing data as indicated previously, this will allow obtaining the most precise model for the variations in Q^n units in Al-free silicate glasses (the case of aluminium deserves its own section, such that its effect will be discussed in section 3). K_{eq}^i values determined for each system are provided in Table 1, based on the data reported in Figure 14.

Metal cation	$K_{eq} \ 2Q^3 = Q^2 + Q^4$	$K_{eq} \ 2Q^2 = Q^1 + Q^3$
Li	0.09(5)	0.2(2)
Na	0.012(4)	0.06(-)

K	0.006(3)	0.012(5)
Rb	0.001(-)	?
Cs	0.001(-)	0.003(-)
Mg	0.464(-)	0.364(-)
Ca	0.106(-)	0.156(-)
Ba	0.078(-)	?

Table 1: mean values of the K_{eq} coefficients for the dissociation of Q^3 or Q^2 units in M^+O_2 -SiO₂ or $M^{2+}O$ -SiO₂ glasses, calculated from the data presented in Figure 14. Standard deviations from the mean are provided in parenthesis, when possible.

Now that we can calculate the fractions of Q^n units for various silicate compositions, it opens doors to calculate the terms of the Adam and Gibbs equation: B_e and $S^{conf}(T_g)$. Le Losq and Neuville (2017) leveraged such ability for K₂O-Na₂O-SiO₂ melts with 60 to 100 mol% SiO₂, taking into account the ideal mixing of Na and K in silicate melts (Richet, 1984) when developing equations for $S^{conf}(T_g)$ and B_e . For $S^{conf}(T_g)$, they assumed that, added to the topological entropy S^{topo} , excess entropies arise from the ideal Si mixing between the different Q^n units S^{mix}_{Si} and random mixing of Na and K into the glass network S^{mix}_{Na-K} , such that:

$$S^{conf}(T_g) = S^{topo} + S^{mix}_{Si} + S^{mix}_{Na-K}, \quad (17)$$

with

$$S^{topo} = \sum_{n=2}^3 x_{Q_{Naenv}^n} S_{Q_{Naenv}^n}^{conf} + \sum_{n=2}^3 x_{Q_{Kenv}^n} S_{Q_{Kenv}^n}^{conf} + x_{Q^4} S_{Q^4}^{conf}, \quad (18)$$

$$S_{Si}^{mix} = \frac{-x_{Si}}{x_O} * 2 * R * \sum_{n=2}^4 x_{Q^n} \ln(x_{Q^n}), \quad (19)$$

and

$$S_{Na-K}^{mix} = \frac{-(x_{Na}+x_K)}{x_O} * 2 * R * (X_K \ln(X_K) + (1 - X_K) * \ln(1 - X_K)), \quad (20)$$

with x_{Si} , x_O , x_{Na} and x_K the atomic fractions of Si, O, Na and K respectively, $X_K = K/(K+Na)$, R the gas constant, and x_{Q^n} the Q^n fractions. We note that, as Le Losq and Neuville (2017) investigated melts with NBO/T lower than 1.5, equations above were limited to Q^2 to Q^4 units. In addition, in the topological contribution (equation 18), Q_{Naenv}^n and Q_{Kenv}^n units with their associated $S_{Q_{Naenv}^n}^{conf}$ and $S_{Q_{Kenv}^n}^{conf}$ values are distinguished; those are the Q^n units, with $n < 4$, that have NBO bonded preferentially to Na or K, respectively. This is an approximation introduced to take into account the fact that Q^n units with NBO

bonded to Na or K may present different S^{conf} values, as shown by the difference of $S^{conf}(T_g)$ between sodic and potassic silicate glasses.

For the calculation of B_e , the link between $S^{conf}(T)$ and B_e needs to be considered. Indeed,

$$B_e = \frac{\Delta\mu s_c}{k_B}, \quad (21)$$

and

$$S^{conf}(T_g) = \frac{N_A}{z(T_g)} s_c, \quad (22)$$

with $\Delta\mu$ the Gibbs free-energy barrier opposed to the cooperative rearrangements of molecular subunits of size $z(T_g)$ and intrinsic entropy s_c , k_B the Boltzmann constant and N_A the Avogadro constant (Adam and Gibbs, 1965; Richet, 1984). s_c represents the entropy of the molecular subunits involved in the melt viscous flow and relaxation process. It appears in both eqs. (21) and (22), yielding a strong correlation between $S^{conf}(T_g)$ and B_e (Neuville and Richet, 1991). From eqs. (20) and (22), we can write

$$B_e = \frac{\Delta\mu z(T_g) S^{conf}(T_g)}{R}, \quad (23)$$

with R the perfect gas constant ($=N_A k_B$). From eq. (23) and considering that $\Delta\mu$ and s_c are temperature-independent (Richet, 1984), B_e should be expressed similarly to $S^{conf}(T_g)$, from the melt structure recorded at T_g in glasses. Considering that, Le Losq and Neuville (2017) assumed that B_e can be expressed as the sum of partial molar B_e terms for the different Q_{Menv}^n and Q^4 units, with additional non-linear contributions arising from the role of $S^{conf}(T_g)$ in eq. (23). The latter non-linear contributions are obtained by directly injecting eqs. (19) and (20) with K_1 and K_2 scaling terms in the expression of B_e :

$$B_e = \sum_{n=2}^3 x_{Q_{Naenv}^n} B_{e_{Q_{Naenv}^n}} + \sum_{n=2}^3 x_{Q_{Kenv}^n} B_{e_{Q_{Kenv}^n}} + x_{Q^4} B_{e_{Q^4}} + K_1 S_{Si}^{mix} + K_2 S_{Na-K}^{mix} \quad (24)$$

Using this model, Le Losq and Neuville (2017) showed that it is possible to calculate the viscosity of the alkali silicate melt with a mean square error of 0.2 log Pa.s. This error is much lower than those of parametric models, typical presenting root mean squared errors on their parametrization datasets of ~ 0.6 - 0.7 log Pa.s (e.g. Hui and Zhang, 2017; Giordano et al., 2008).

Le Losq and Neuville (2017) used cubic splines for interpolation of the Q^n distribution in their model. With the approach presented in this chapter, it is possible to replace this by the K_{eq} -based model for Q^n distribution in silicate melts in order to improve the future deployment and development of such a thermodynamic viscosity model. Preliminary tests yield good predictions with standard errors of 0.3 log

Pa·s, a value slightly higher than the initial value of 0.2 log Pa·s from Le Losq and Neuville (2017). This difference arises from the fact that the initial cubic spline modeling of the Q^n distribution in K₂O-Na₂O-SiO₂ melts is slightly more accurate than the global structural model based on the K_{eq} values calculations. Differences are in the range of 1-2%. This actually highlights how sensitive a model can be to melt structure: an error of only 1-2 % alters measurable the accuracy of the model.

3. Aluminosilicate glasses and melts

Aluminum is a key element in Earth and material sciences because it is present, in various amounts, in most geologic and industrial glasses, and it drastically affects their properties. Generally, a well-known feature of the structure of silicate minerals is that silicon ions may be replaced isomorphously by aluminum ions provided that either mono or divalent cations are available to maintain local neutralization of electrical charges in the structure (Putnis, 1992). This feature is clearly visible in tectosilicates that present an interconnected three-dimensional silica-like tetrahedral structure, in which the mono or divalent cations are accommodated in the framework free volumes. In these minerals, the ratio $R = M^+_2O/Al_2O_3$ or $M^{2+}O/Al_2O_3$ is unity. The metal cations play charge compensator roles, and no non-bridging oxygens are present. This is also generally the case in tectosilicate glasses, albeit some possible minor amounts of NBOs may be present (Stebbins and Xu, 1997). However, an additional complexity in aluminosilicate glass is the variable coordination number of Al. Indeed, while Al mostly substitutes for Si in the SiO₄ tetrahedra in silicate glass, it also is found in five and six-fold coordinations, in proportions that depend on melt chemical composition, system pressure and temperature, and glass-forming cooling rate (Ohtani et al., 1985; Risbud et al., 1987; Murdoch et al., 1985; Mc Millan, et al., 1982; Merzbacher et al., 1990, 1991; Lacy, 1963; Merzbacher and White, 1991; Coté et al., 1992; Stebbins and Xu, 1997, Sato et al., 1991; Allward and Stebbins, 2005a,b, 2007, Allwardt, et al., 2004; 2007; Stebbins et al., 2000, 2008; Stebbins, 2008; Henderson et al., 2007; Lee 2005, Lee and Stebbins, 2000, 2002, 2009; Neuville et al., 2004ab, 2006, 2008a,b, 2010; Le Losq et al., 2014). The varying coordination number of Al raises the question of its role in the melt structure depending on its coordination. Indeed, following the terminology proposed by Zachariassen (1932), while Al in four-fold coordination is a network former, Al in six-fold coordination should be a network modifier, and Al in five-fold coordination should probably also be considered as a network modifier. We will review such questions, and more generally the connection between aluminosilicate melt structure and properties in this section.

3.1 Molar volume, coordination number and structure

Density and molar volume are properties easy to measure in glasses, and bring important pieces of information regarding their internal structure, and by extension, that of the melt at the glass transition.

Variations in the molar volumes of alkali and alkaline-earth silicate and aluminosilicate glasses as a function of silica concentration show similar trends, except for the heaviest elements, like potassium, cesium and rubidium (Figure 15). From silica glass with $V_m=27,31\text{cm}^3$, it is found that generally the molar volume decreases with the decrease in SiO_2 , except for K-, Cs-, Rb- silicate glasses. For the same SiO_2 content, the magnitude of the decrease of the molar volume depends on the ionic field strength of the metal cation: it is moderate for Na^+ and very important for Mg^{2+} (for example V_m of Ca silicate $=21,21\text{cm}^3$ and $24,14\text{cm}^3$ for Na-silicate, both at 60 mol % of SiO_2 – note that we compare Ca and Na silicate glass and not Mg-silicate at 60 of silica, because this glass is made with a high cooling rate, and its fictive temperature is probably very different than that for the Na-silicate glass). In the case of potassium, V_m always increases regardless of the glass aluminum content. For other cations, V_m variations in aluminosilicate glasses are more complex. First, the M^{2+}/Al or M^+/Al ratio directly influences the variations of V_m with the glass silica concentration. Secondly, a general concave trend of the V_m versus SiO_2 relationship is observed according to previous work (Seifert et al., 1982, Doweidar et al., 1999, Doweidar 1999). Taking the SiO_2 - CaAl_2O_4 binary mixture as an example, and starting from silica, V_m decreases with silica decrease until 50 mol% SiO_2 ($V_m=25,81\text{cm}^3$ for anorthite glass), and then increases up to a value of $\sim 27\text{cm}^3$ for the pure CaAl_2O_4 end-member for example. Such non-linear V_m trends indicate that V_m modelling in aluminosilicate glasses cannot be obtained from a linear combination of partial V_m values assigned to the different oxides. To circumvent this problem, several papers have modelled the density of glass and melts as a function of their expected Q^n speciation (Doweidar, 1999, Doweidar et al., 1999) or other entities like partial volume of oxide components (Lange, 1996, 1997, Bottinga and Weill, 1970).

One of the most important pieces of information arising from V_m data analysis is the possibility to provide knowledge regarding the coordination numbers of metal cations as well as aluminum. We have mentioned, in section 2.1 (Figure 5), that metal cations' CNs remain almost constant in silicate glasses. It is also well known that the molar volume of the SiO_4^{4-} tetrahedra doesn't change a lot as a function of the Q^n species because the $d(\text{Si-BO})$ or $d(\text{Si-NBO})$ vary less than 0.05\AA (Cormack and Du, 2001; Du and Cormack 2004). This variation corresponds to a variation of around 3% on the distance between Si-BO and Si-NBO which causes a variation of the volume of the tetrahedra between 12.4 and 9.3% as a function of the amount of alkali content respectively 10 and 50%, if we used the distances given by Cormack and collaborators. In the case of aluminum, the molar volume of AlO_4^{5-} tetrahedra doesn't change because Al-O distances stay almost constant as a function of the different Q^n species for Al^{IV} , and the most important change in the molar volume of Al units come from the change in the coordination of Al and

⁷ The distance Al-O increases from 1.76 up to 1.86 Å when the coordination number of Al changes from 4 to 6 (Cormier et al., 2000, 2003, Hennet et al., 2016). This implies that Al-O distance vary smaller when Al change of Q^n species in tetrahedra.

not of the Q^n species of Al. If we consider that the volumes vary very slightly in the ternary diagrams this implies that the molar volume variation observed between silicates and aluminosilicates glasses is produced essentially by changes in coordination number of alkaline or alkaline-earth. Before discussing this, we will quantify the different proportions of aluminum.

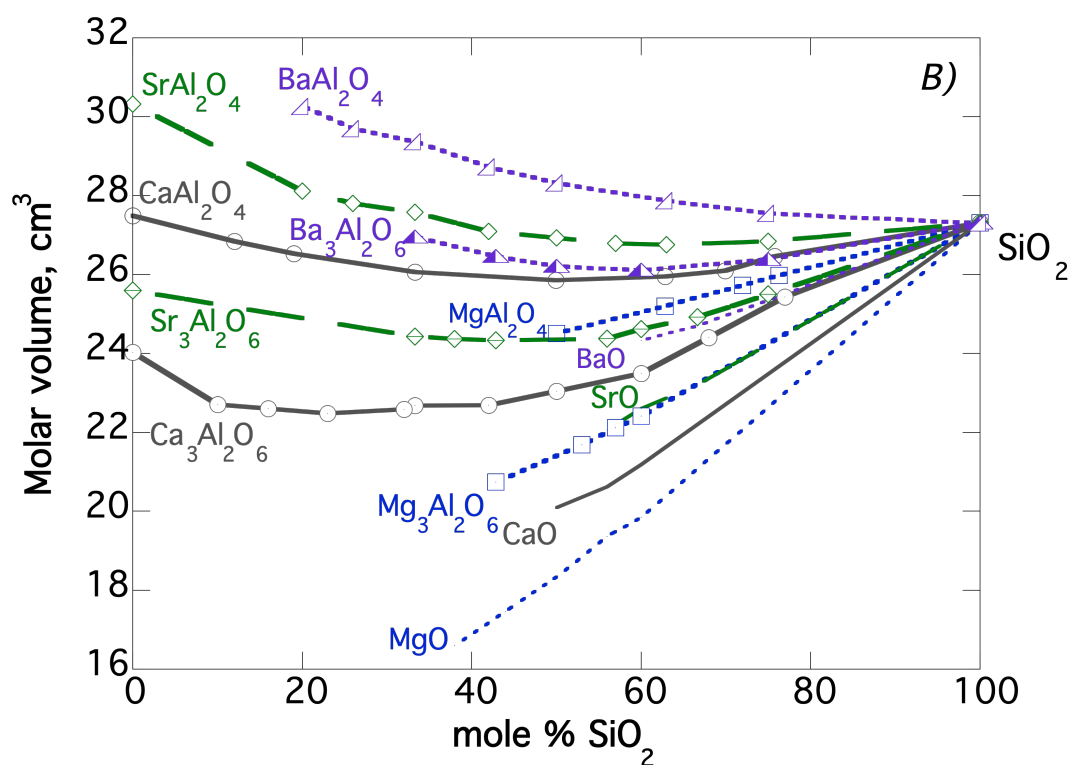
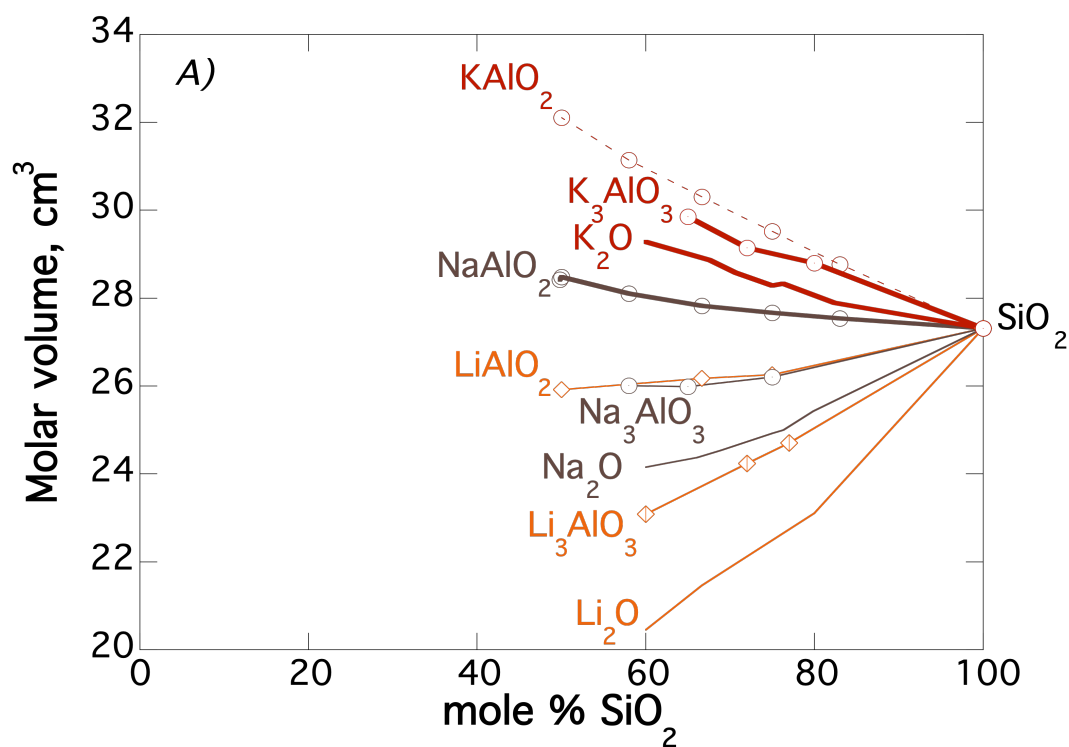


Figure 15: A) molar volume for alkali and B) for alkaline-earth silicate and Al-silicate glasses, data for NAS, KAS compositions (Le Losq et al. 2017, 2015, Neuville, 1991, 2006), LAS (Strukelj, 2008), MAS, CAS system (Neuville, 1992), SAS data from Novikov et al., (2017), BAS original data and from Novikov (2017).

3.2 Proportion of Al in five-fold coordination

Given the role of Al coordination on V_m and other melt/glass properties, it is clear that Al coordination should be quantified. Fortunately, through ^{27}Al NMR spectroscopy, several studies allow drawing a picture of Al coordination variations with composition, temperature and pressure.

The presence of $^{[5]}\text{Al}$ is systematically observed in the middle of the ternary diagrams of Li aluminosilicate, and alkaline-earths aluminosilicate, including peralkaline glasses in which Al in five-fold coordination is not expected considering the stoichiometry (Neuville et al., 2006, 2007, 2008b, 2010). Moreover, the proportion of $^{[5]}\text{Al}$ is always maximum for glasses with about 50 mol % of SiO_2 (Neuville et al. 2004, 2006, 2008, Novikov et al., 2017). On the contrary, Na or K tectosilicate and peraluminous glasses have small concentrations of Al in five-fold coordination, which becomes undetectable in peralkaline compositions (Stebbins and Farnan, 1992; Allwardt et al., 2005a,b; Lee et al. 2009; Le Losq et al., 2014). The variation of NMR chemical shift values for $^{[5]}\text{Al}$ and $^{[6]}\text{Al}$ sites with SiO_2 content are less dependent on silica content than $^{[4]}\text{Al}$ sites (Figure 13b, from the previous chapter). The NMR chemical shift values are directly related to the inter-tetrahedral angle or the first-neighbor cations (Stebbins and Farnan, 1992). This can imply that $^{[5]}\text{Al}$ and $^{[6]}\text{Al}$ sites are more connected with Al in four fold coordination than with Si in different Q^n species which does not respect the Lowenstein rule's, but in the majority of studies this rule is avoided (Neuville et al., 2008a, 2010, Lee and Stebbins, 2009, Hiet et al., 2009). 2D NMR ^{27}Al - ^{29}Si correlations (Lee and Stebbins, 2009, Hiet et al., 2009) show that $^{[5]}\text{Al}$ is more connected with Si than with Al in lanthanum aluminosilicate glasses (Florian et al., 2007). Highly coordinate Al species, especially $^{[5]}\text{Al}$, are likely to have a positive formation enthalpy and their abundance in the melt is expected to increase with increasing temperature as observed by some authors (Neuville et al., 2008b; Stebbins, 2008). However, the ^{27}Al NMR spectra at high temperature shifts to lower ^{27}Al chemical shift values with increasing temperature, this can be interpreted as resulting from the appearance of $^{[5]}\text{Al}$ and $^{[6]}\text{Al}$ contributions located at lower chemical shifts than that of $^{[4]}\text{Al}$ (Kanehashi and Stebbins, 2007, Thompson and Stebbins, 2011, 2012, 2013). Signatures of Al in high coordination are also clearly visible using X-ray absorption spectroscopy at the Al K-edge (Neuville et al., 2008b). In the case of CA50.25 (anorthite melts composition), Neuville et al (2008b) show an increase from 8% up to 22% of $^{[5]}\text{Al}$ between room temperature and 1900 K, in good agreement with high

temperature Al NMR (Stebbins, 2008). A significantly higher proportion of $^{[5]}Al$ is observed in magnesium aluminosilicate glasses than in the others aluminosilicate systems on the tectosilicate join (Neuvillle et al., 2008b, Novikov et al., 2017, Figure 23B) which can be explained by the promotion of highly coordinated Al species with the presence of metal cations with relatively high ionic field strength, as shown in boroaluminate by Bunker et al., (1991) and in La-aluminosilicate glass compositions (Florian et al. 2007). Such an idea agrees with the increase in the disorder when the cation field strength decreases (Mysen et al., 1982; Murdoch et al., 1985; Sharma et al, 1985; Navrotsky et al., 1982, 1985; Lee and Stebbins 2000, 2002, Neuvillle et al. 2008a, Wu and Stebbins, 2009; Lee et al., 2016, Le Losq et al., 2017).

Several papers on ^{27}Al NMR spectroscopy have shown that, for $R = M/Al$ varying between infinity and 1 (silicate to tectosilicate compositions), Al is essentially in four-fold coordination (Risbud et al., 1987, Murdoch et al., 1985; Merzbacher et al., 1990, 1991; Merzbacher and White, 1991; Coté et al., 1992, Stebbins and Xu, 1997, Stebbins et al., 2000, 2008; Neuvillle et al., 2004ab, 2006, 2007, 2008a, McMillan and Kirkpatrick, 1992; Massiot et al., 2008; Lee and Stebbins, 2009, 2016; Richet et al., 2009a; Le Losq et al., 2014; Novikov et al., 2017; Florian et al., 2018, Charpentier et al., 2018). In the CAS system, 93 % of Al is in 4-fold coordination, and the proportion of $^{[5]}Al$ increases slowly up to 7% for the anorthite glass composition on the tectosilicate join (Neuvillle, et al., 2006). In the MAS system, the proportion of $^{[5]}Al$ is a little higher, 12% for Mg-anorthite glass (Neuvillle et al., 2008a). For all others M^+ or M^{2+} Al-silicate systems, the proportion of $^{[5]}Al$ is lower and the amounts decrease with increasing z/r (z : cation charge and r atomic radius, see Neuvillle et al., 2008a or Stebbins et al., 2008, Novikov et al., 2017). Figure 16 A and B show the total amount of $^{[5]}Al$ in mol % in the CAS and MAS system.

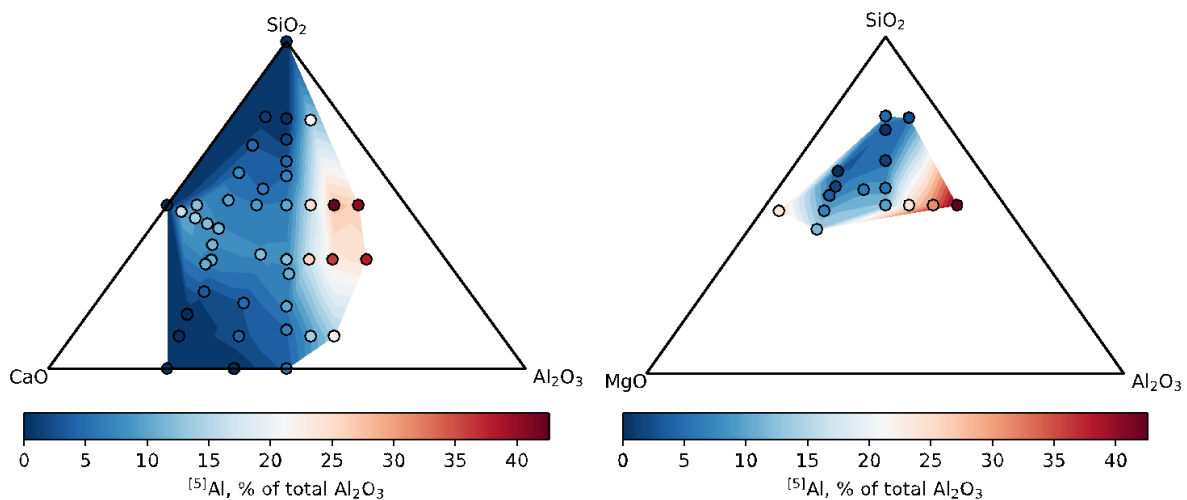


Figure 16: Percentage of $^{[5]}Al$ of total Al_2O_3 in the CAS (left) and MAS (right) systems. Data from Neuvillle et al. (2004, 2006, 2007, 2008b, 2010, Massiot et al., 2008; Licheron et al., 2011).

3.3 Molar volume, AlO_4^{4-} and implications for the coordination number of metal cations

In the previous sections, we saw that the molar volume of Si in the different Q^n species can vary by less than 10% when n varies from 4 to 0 and that the molar volume of Al stays almost constant for Al in four-fold tetrahedra. Despite this, the molar volume V_m of a lime aluminosilicate glass with 50% of SiO_2 (wollastonite to anorthite glasses) can vary from 20 to 25.9 cm³ between the wollastonite and anorthite compositions, this represents an increase of 25 %. This increase is 12 % between silicate glasses and the glasses on the R = 3 join, and of 12 % again between the join R = 3 and the join R = 1 with R=MO/Al₂O₃ with M=Mg, Ca. This huge increases of V_m is a maximum at 50 mol % silica.

In the case of calcium aluminate glasses, the molar volume difference between C3A (Ca₃Al₂O₆) and CA (CaAl₂O₄) glasses is only 6.8 % (Figure 15B). In Figure 15B, the increase of molar volume between join R=3 and 1 can also result from a change in the coordination number of calcium because the aluminum stays in four-fold coordination between C3A and CA glasses. In this case, Ca varies between regular octahedra where Ca is in 6-fold coordination to a distorted site in 7-8-fold coordination (Cormier et al., 2001, 2003, Neuville et al., 2008b, 2010, Cicconi et al., 2016; Drewitt et al., 2012, 2017, Jakse et al., 2012; Takahashi et al., 2015; Hennet et al., 2016).

Generally, the coordination number of alkali or alkaline-earth elements is difficult to determine, but changes in the V_m of a glass can inform us about CN changes. For instance, the above variation in the CN of Ca is in good agreement with the changes observed via XANES at the Ca *K*-edge in Ca aluminosilicate glasses (Neuville et al., 2004b, 2008, Cormier and Neuville, 2004; Cicconi et al., 2016). In particular, Cicconi et al. (2016) observed a shift to greater energy and a decrease in intensity of the pre-edge in the XANES spectra at the Ca *K*-edge between wollastonite to anorthite glasses. This shift can also be correlated with the new peak ω_2 observed in the depolarized Raman spectra, VH, (see Figure 13, chapter 3 this volume, Drewitt et al., 2021). The peak ω_2 decreases when R goes from infinity to 1 (to silicate to tectosilicate glasses)

In the case of the NAS system (Na₂O-Al₂O₃-SiO₂), the peak ω_2 is also clearly visible for silicate glasses in the VH Raman spectra and it disappears for R=Na₂O/Al₂O₃=1 (Hehlen and Neuville, 2015). This peak ω_2 exists in all silicate glasses with M⁺=Na, K and M²⁺=Mg, Ca, Sr, Ba (Hehlen and Neuville, 2015, 2020) and its intensity decreases when Al replaces M⁺ or M²⁺. In the NAS glass system, a shift is also observed in the chemical shift, δ , of ²³Na NMR between NS3 and albite glass (Le Losq et al., 2015) following a trend that can be correlated with variations in XANES at the Na *K*-edge between the NS3 and albite glasses (Neuville et al., 2004).

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In the case of the SAS (SrO-Al₂O₃-SiO₂) system, Novikov et al. (2017) also observed some change in the first XANES oscillation which can also be attributed to a change in CN of strontium.

To summarize, ²³Na NMR, XANES at the Na, Sr, Ca *K*-edge, and depolarized Raman spectra, (VH), all show changes that can be correlated with an observed increase in the molar volume between silicate and tectosilicate glasses. *V_m* varies between 4% and 16% with changing the M⁺/Al or M²⁺/Al ratio at a given silica content and depends clearly on the metal cation size. The *V_m* changes can be assigned to changes in the coordination of metal cations as their role in the aluminosilicate network varies with the M/Al ratio. Similarly, features observed in NMR, XANES or Raman spectroscopies can be directly correlated with such change in the role of alkali or alkaline-earth elements between network modifier and charge compensator. The concept of network modifier and charge compensator was introduced by Zachariazen but never really demonstrated. Now, with the technical evolution of spectroscopy methods, but with measuring the molar volume determined from density measurements, it is possible to have an idea of the role alkaline elements.

In Figure 17, the molar volume of various glasses is shown as a function of mole percent of Al₂O₃ at constant silica. The glass compositions vary between MSiO₃ up to Al₂SiO₅, M=Mg, Ca, Sr, Ba. Four trends are clearly visible for the four alkaline-earth aluminosilicate glass families. In the case of MgSiO₃-Al₂SiO₃, the molar volume varies almost linearly with Al₂O₃ (the linear variation corresponds to the dashed line). This can imply:

- i) that the coordination number of Mg is almost constant in this system and is not really affected by the Mg/Al substitution, an idea in good agreement with XANES at the Mg *K*-edge (Trcera et al., 2009);
- ii) that the molar volume depends on the proportion of Al in five-fold coordination.

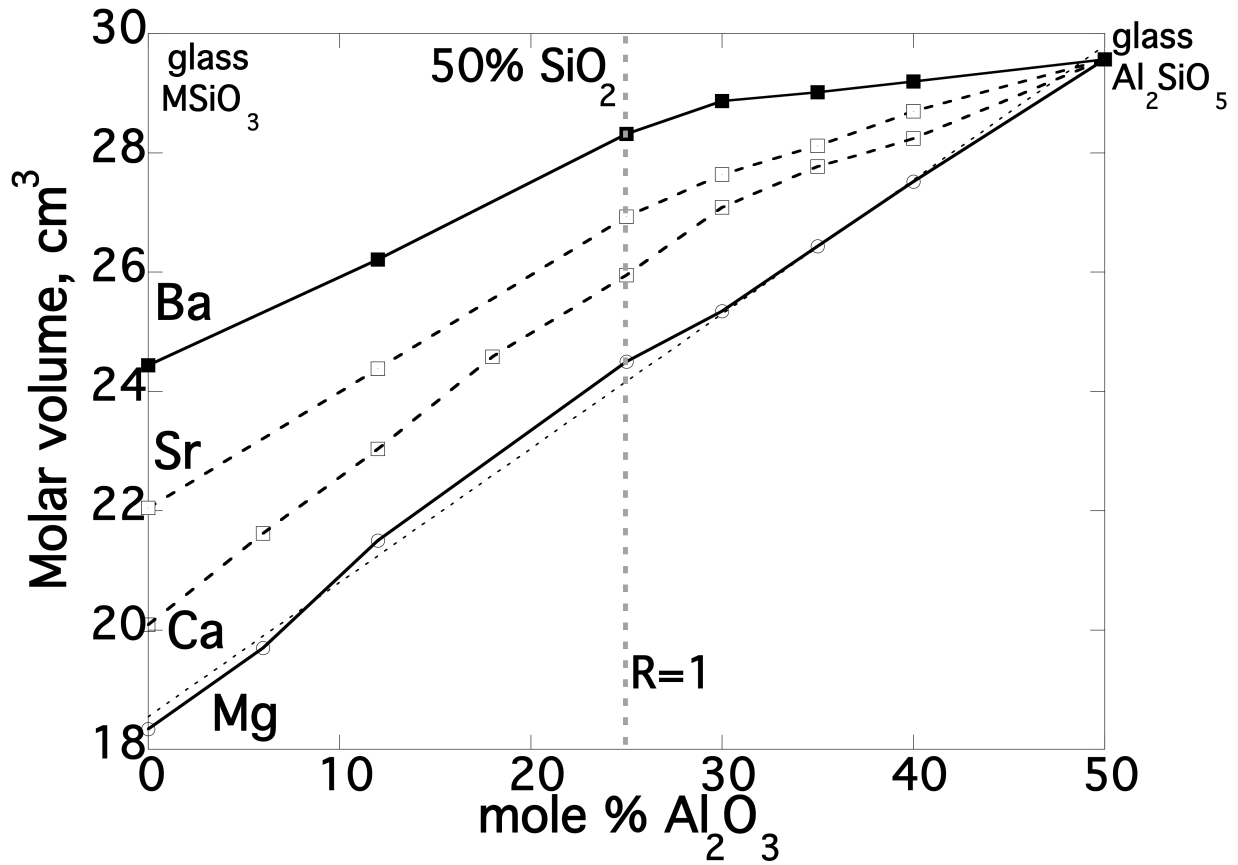


Figure 17: molar volume at constant $\text{SiO}_2 = 50$ mol% for four alkaline-earth aluminosilicate glass compositions. (The dashed line between MgSiO_3 and Al_2SiO_5 glasses correspond to this equation $18.547 + 0.22497 \cdot \text{Al}_2\text{O}_3$). Values are calculated from density measurements from Neuville (1992) for Ca and Mg system peralkaline compositions, Neuville et al. (2006, 2008) for Ca and Mg Per-aluminate compositions, Novikov et al. (2017) for Sr compositions, Novikov (2017) for Ba compositions and Al_2SiO_5 are from Wang et al. (2020).

For the Ca, Sr and Ba aluminosilicate glasses the changes observed in the slopes near 25 mol % of Al_2O_3 correspond to the tectosilicate line ($R = M^{2+}O / \text{Al}_2\text{O}_3 = 1$), characterized by an increase in the proportion of $[\text{Al}]$ in the glass (Neuville et al., 2006, 2008, Novikov et al., 2017, Novikov 2017). This Figure clearly shows two domains, one where V_m is affected by aluminum, the peraluminous domains, and one where V_m depends on an increase in the coordination of the alkaline-earth element, the peralkaline domain (Cicconi et al., 2016; Novikov et al., 2017).

3.3 Glass transition temperature

In general, a decrease in T_g occurs with decreasing SiO_2 content at all R ratios ($R = M^{2+}O / \text{Al}_2\text{O}_3$ or $M^{2+}O / \text{Al}_2\text{O}_3$) (Figure 18). It is particularly important at silica contents between 100 and ~75 mol%. In

detail, this T_g decrease is more significant in Al-free silicate melts than in tectosilicate glasses; the lower R , the lower the effect of changing SiO_2 on T_g . Typically, at $R=1$ in tectosilicate compositions, the decrease in T_g is of ~ 300 K for $\text{MAI}_2\text{Si}_2\text{O}_8$ glasses, and 200 K and 300 K for respectively KAlSiO_4 and NaAlSiO_4 glasses. The large contrast between T_g variations in Al-free silicate and tectosilicate glass compositions results from the changing role of metal cations upon aluminum addition into the melt network. Indeed, while in silicate glasses, alkali and alkaline-earth elements are network modifiers that break the 3D tetrahedral network, in aluminosilicate glasses they ensure the charge compensation of AlO_4^- tetrahedral units and thus participate in increasing melt polymerization. This translates to higher T_g along the $\text{M}^+\text{AlO}_2\text{-SiO}_2$ or $\text{M}^{2+}\text{Al}_2\text{O}_4\text{-SiO}_2$ binaries than along the $\text{M}^+\text{O}_2\text{-SiO}_2$ or $\text{M}^{2+}\text{O-SiO}_2$ binaries.

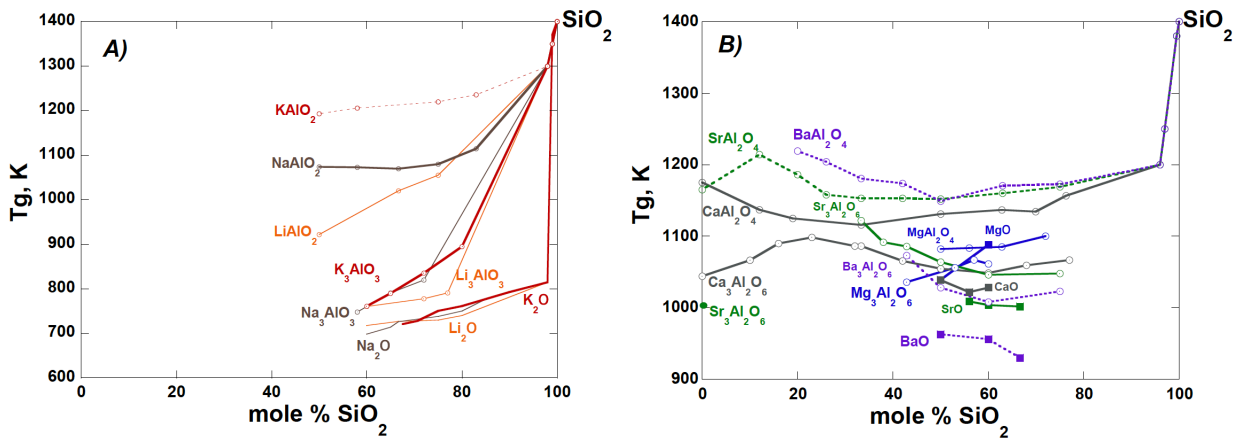


Figure 18: Glass transition temperature T_g , in K, of silicate and aluminosilicate glasses with alkali elements (A) and alkaline-earth elements (B). The glass transition temperatures were determined from viscosity measurements, i.e. they correspond to the temperature at which $\log \eta = 12$ Pas. Data from Neuville (1991) for Li, K silicates; Neuville (2006) and Le Losq et al. (2014) for sodic systems, (Neuville, 1992) for Ca and Mg, Novikov et al. (2017) for Sr and Novikov (2017) for Ba.

3.4 Link between structure and properties of aluminosilicate melts: example of the $\text{CaO-Al}_2\text{O}_3\text{-SiO}_2$ system

To discuss and illustrate the links between glass structure and glass/melt properties, the $\text{CaO-Al}_2\text{O}_3\text{-SiO}_2$ (CAS) glass system will be taken as an example in the following. The CAS system has the advantage having the largest glass forming domain (Neuville et al., 2006). The observations and interpretations made in this system can be extended to other ternary system. In general, in the CAS system, an important decrease in T_g is observed with decreasing SiO_2 content, which results from network depolymerization, except at $R = 1$. In the latter case, melt polymerization is almost constant and the decrease in T_g results from the formation of weaker Si-O-Al bonds as Al substitutes for Si in Q^4 units (Navrotsky et al., 1985;

299 [Neuvville et al., 2004](#), [Cormier et al., 2001, 2005](#), [Hennet et al., 2016](#)). T_g varies by about four hundred
 300 degrees between pure SiO_2 ($T_g \sim 1480$ K; [Richet and Bottinga 1984](#)) and the center of the ternary system
 301 ($T_g = 1115$ K for $\text{CA}33.33$, $\text{M}^{2+}\text{AX.Y}$ where M^{2+} =alkaline-earth element and $\text{A}=\text{Al}_2\text{O}_3$, $\text{X} = \text{SiO}_2\%$ i.e.
 302 33.3 mol%, $\text{Y} = \%\text{Al}_2\text{O}_3$ and $\text{M}^{2+}\text{O}=(100-(\text{X}+\text{Y}))$ in this case, $\text{CaO}=33.3\%$). At $R=1$, glasses with less
 303 than 30 mol% SiO_2 display a continuous increase in T_g upon decreasing SiO_2 content, and T_g reaches a
 304 new local maximum for the CaAl_2O_4 composition at 1175 K; the minimum in T_g at $R=1$ is centered
 305 between 40 and 50 mol% SiO_2 . The observed behavior in T_g at $R = 1$ results from the competition between
 306 Si and Al as network formers. As indicated by the difficulty of forming glasses at high Al concentrations
 307 (e.g. along the SiO_2 - Al_2O_3 binary) as well as a lowering of T_g with addition of Al. Al can be considered
 308 as a relatively poor network former compared to Si. This is further corroborated by variations in the melt
 309 fragility (i.e. the derivative of the melt viscosity as a function of T at T_g): the strongest existing melt is
 310 SiO_2 , and Al-rich melts have higher fragility, with the most fragile being CaAl_2O_4 in the CAS system
 311 (Figure 19). This is also in good agreement with the variation of the C_p^{conf} of SiO_2 and CaAl_2O_4 , which
 312 varies between 8.2 to 13.8 J/mol K. This ΔC_p between liquid and glass at T_g can be expressed in term of
 313 C_{pl}/C_{pg} ratio as proposed by Angel ([1991](#)). This ratio varies between 1.1 for SiO_2 , the strongest melt to
 314 1.37 for CaAl_2O_4 , the more fragile melt in Figure 19.

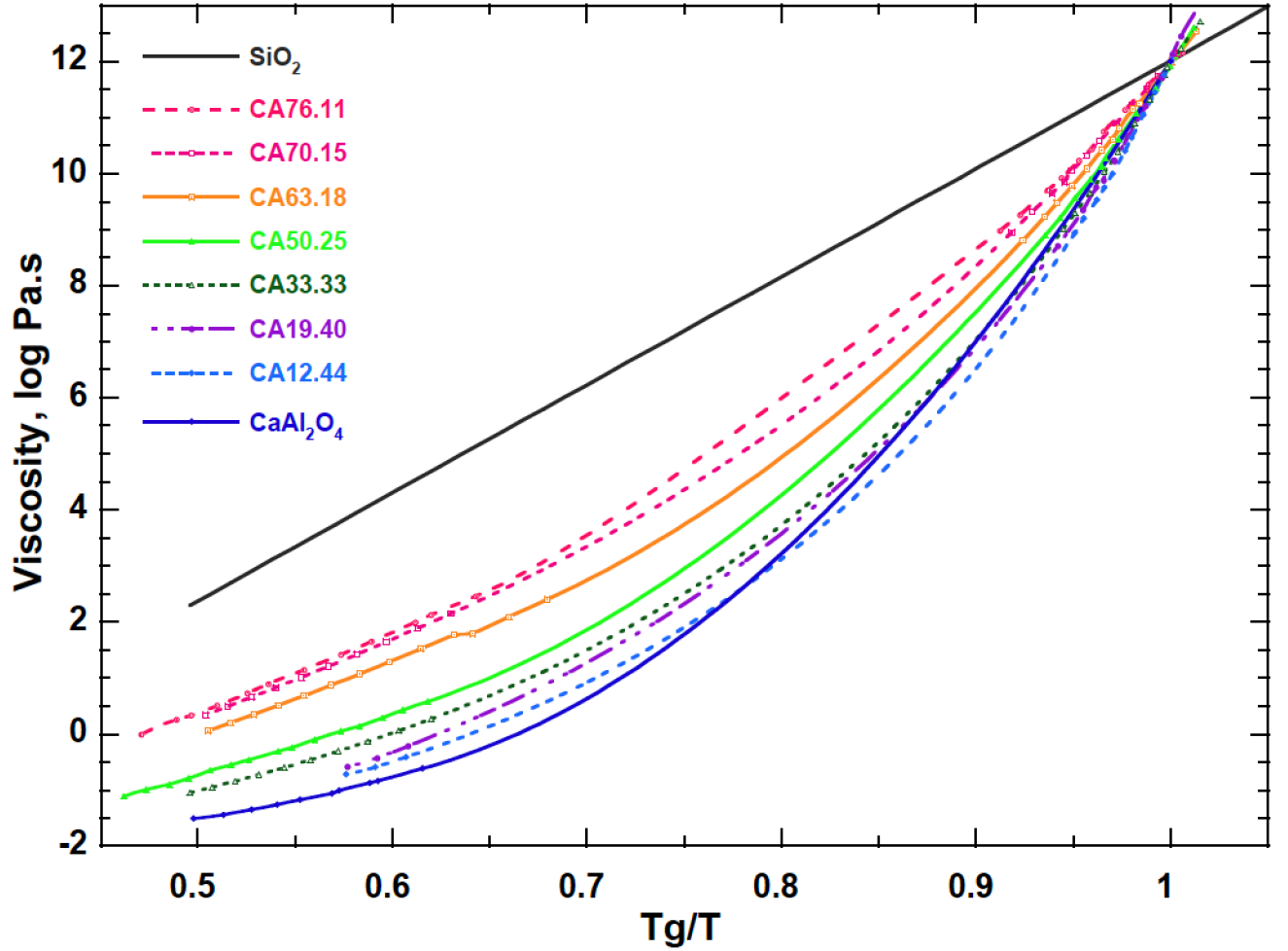
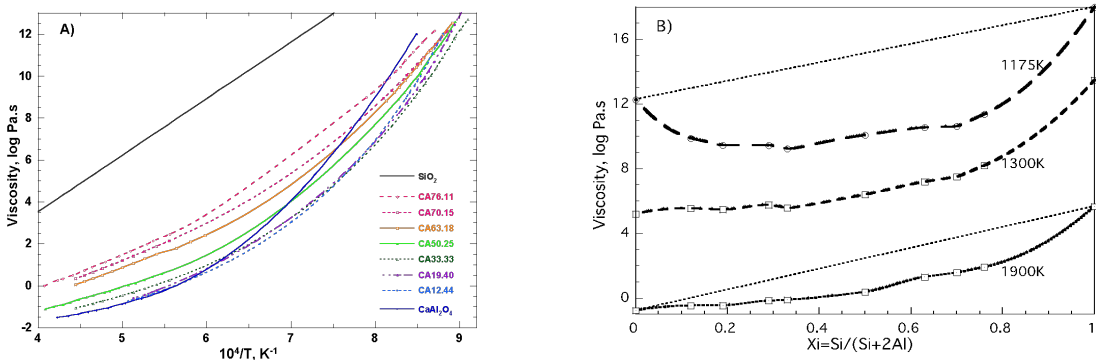


Figure 19: “Angel plot” showing $\log \eta$ versus T_g/T for melts with $R=\text{CaO}/\text{Al}_2\text{O}_3=1$ in the CAS system.

As mentioned previously, the CAS system presents one of the largest glass-forming domains. Indeed, it is possible to quench glasses even without silica, e.g. like the C12A7 composition (CXAY with C=CaO, A=Al₂O₃, X and Y are their relative proportions), using classic quench rates. With increasing quench rates, e.g. via laser levitation heating, it is possible to make glasses between C3A and CA1.5 (from 75 down to 40 mol% CaO). In this condition, the addition of a small amount of SiO₂ in CaO-Al₂O₃ glasses extends the glass forming ability and lowers the melt liquidus temperature. Other macroscopic properties are also markedly changed with silica introduction in CaO-Al₂O₃ melts/glasses, suggesting extensive structural modifications. For instance, an anomalous behavior of T_g is clearly observed at low silica content, with a maximum in T_g visible at ~20 mol% SiO₂ along the Ca₃Al₂O₆-SiO₂ binary (Figure 18B; Higby et al., 1990; Neuville 1992; Neuville et al., 2004; Cormier et al., 2005; Neuville et al., 2010). Note that this anomalous behavior of T_g results from the change in the polymerization of Al which varies between Q^2 to Q^4 species along the join $R=3$ between C3A and CA20.40 glasses (Neuville et al., 2004, Cormier et al., 2001, 2003, 2005).

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To explore further the properties of CAS glass compositions, viscosity along the join R=1 (CaAl₂O₄ – SiO₂ binary) is plotted in Figure 20. At high temperatures, viscosity decreases constantly with addition of CaAl₂O₄ (Figure 20B). At supercooled temperatures, near T_g , melt viscosity decreases by ~ 7 orders of magnitude with adding 30 mol% CaAl₂O₄ to silica, then a further addition of 30 mol% CaAl₂O₄ yields a decrease in viscosity of 1 order of magnitude. Between 40 and 0 mol% SiO₂, melt viscosity increases ~ 5 orders of magnitude. The deviation from a linear variation (Figure 20B) is 4 orders of magnitude at 1175 K, and less than 2 orders of magnitude at 1900 K. It thus decreases with increasing temperature. These changes thus deviate strongly and negatively from a linear variation, particularly at undercooled conditions. Such behavior may originate from mixing between Si and Al in Q^4 units that occurs in such tectosilicate melts (Seifert et al., 1982), this mixing resulting in excess entropy and leading to non-linear, convex variation of the melt viscosity as a function of the $Si/(Si+Al)$ ratio (e.g., Neuville and Mysen, 1996).



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Figure 20: A) Viscosity of melts between silica and CaAl₂O₄ compositions, CAX.Y mean: X=SiO₂ %, Y=Al₂O₃, CaO=100-(X+Y). B) viscosity at constant temperature, 1150, 1200, 1800 K as a function of $Xi=Si/(Si+2Al)$. SiO₂ data are from Hetherington et al. (1964) Leko, (1979) and Urbain et al. (1982), CaAl₂O₄ from Urbain (1983), Neuville (1992).

In all others glass systems, alkali or alkaline-earth aluminosilicate, viscosity shows similar behavior along the join $R=1$, an important decrease between 100 and 70% of SiO₂, and for lower silica content, a smaller decrease of the viscosity (Alkaline see Le Losq et al., 2017, Mg, Neuville 1992, Sr, Novikov et al 2017, Ba, Novikov 2017).

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As already mentioned, the configurational heat capacity, Cp^{conf} , is required in order to determine the configurational entropy, $S^{conf}(T)$, from equations (6) and (7). In the CaO-Al₂O₃-SiO₂ system, the

configurational heat capacity can be determined from the difference between the heat capacity of the melt at T and that of the glass at T_g (equation 8). Several models exist, like those of Richet (1987) and Richet and Bottinga (1985), and allow one to calculate the configurational heat capacity with precision. The more recent works of Courtial (1993) and Richet and Neuville (1992) allow refining the heat capacity modelling in the CaO-Al₂O₃-SiO₂ system, with using the following partial molar heat capacity equations (J mol⁻¹ K⁻¹) for the glass and liquid (the estimated uncertainties is about 1% for the partial molar heat capacities following Courtial and Richet, 1993):

$$C_{pg\ SiO_2} = 127.2 - 0.010777T - 431270/T^2 - 1463.9/T^{0.5}, \quad (25)$$

$$C_{pg\ Al_2O_3} = 175.46 - 0.005839T - 1347000/T^2 - 1370/T^{0.5}, \quad (26)$$

$$C_{pg\ CaO} = 39.159 + 0.01865T - 152300/T^2, \quad (27)$$

$$C_{pl\ SiO_2} = 81.37, \quad (28)$$

$$C_{pl\ Al_2O_3} = 85.78 - 130.216T, \quad (29)$$

$$C_{pl\ CaO} = 86.05. \quad (30)$$

With these partial molar heat capacity terms, $C_p^{conf}(T_g)$ can be calculated easily and used for the calculation of $S^{conf}(T_g)$, which are plotted in Figure 21. It is clearly visible that $S^{conf}(T_g)$ of silica glass is almost the same than that of CaAl₂O₄, CA glass. Assuming that $S^{conf}(T_g)$ can be decomposed into topological and excess terms, we can model the observed variations as

$$S^{topo} = X_{SiO_2} S^{conf}(T_g)_{SiO_2} + (1 - X_{SiO_2}) S^{conf}(T_g)_{CaAl_2O_4}, \quad (31)$$

S^{topo} is thus assumed to vary linearly between the two end-members SiO₂ and CaAl₂O₄ glasses. This term S^{topo} , is independent of temperature and corresponds to a simple mechanical mixing between two entities. The term S^{ideal} , defined in equation 10, can be taken as equal to the entropy obtained from an ideal random mixing of two species, in this case Si and Al. At high or low concentration of silica, $S^{conf}(T_g)$ variations seem to follow the ones predicted assuming such an ideal mixing between Si and Al. This suggests that CaAl₂O₄ randomly replace Si₂O₄ in the melt/glass structure. However, between 33 and 75 percent of silica, $S^{conf}(T_g)$ variations depart from the predictions of this ideal mixing model. In this silica concentration range, a decrease of ~25 % in $S^{conf}(T_g)$ is observed, and there is a maximum difference of ~2.5 J mol⁻¹ K⁻¹ between $S^{conf}(T_g)$ values and the ideal model at around 50 mol% SiO₂ (Figure 21).

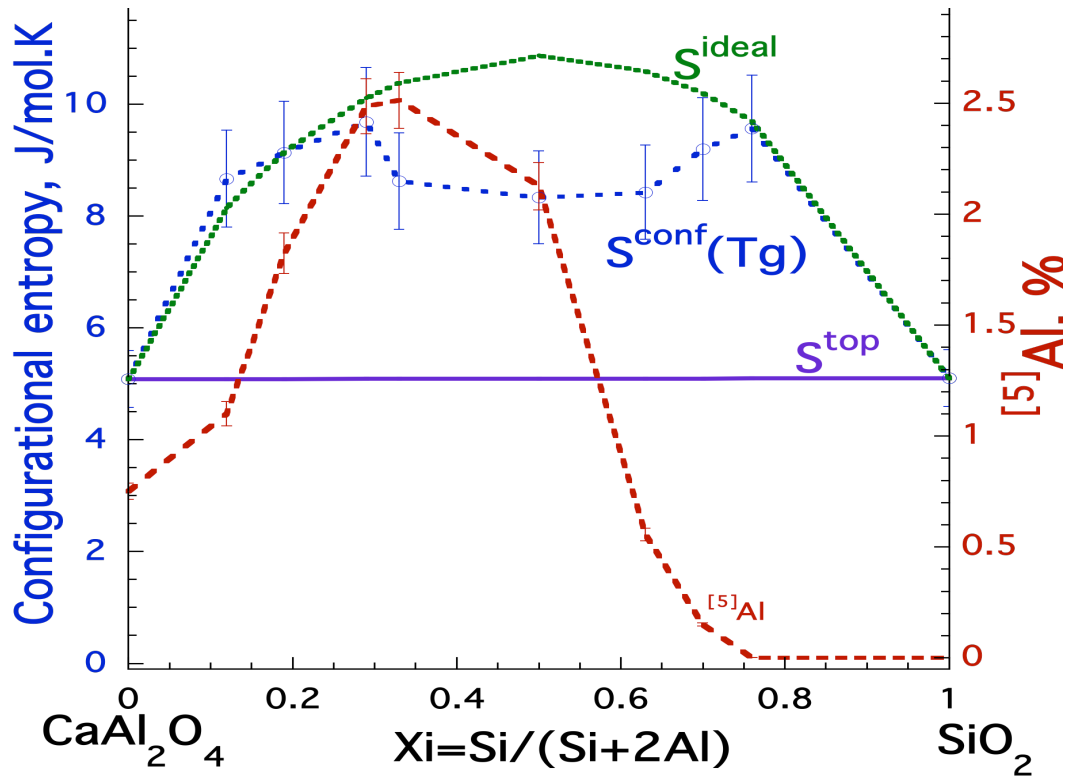


Figure 21: Configurational entropy at the glass transition temperature, $S^{conf}(T_g)$ (Neuvill, 1992), and real proportion of $[5]Al$ in mole percent in glasses along the SiO_2 - $CaAl_2O_4$ binary (Neuvill et al., 2004, 2004). Each $S^{conf}(T_g)$ value was calculated from the viscosity data obtained for each composition.

The observed $S^{conf}(T_g)$ variations (Figure 21) suggest that :

- i) at low or high silica concentrations, Si and $[4]Al$ mix almost randomly in the melt/glass structure,
- ii) a deviation from ideal mixing exists at intermediate silica contents. The latter deviation can be rationalized upon considering that Al in five-fold coordination represents around 2,5 % of total Al in the CA50.25 glass.

Excess $[5]Al$ tends to decrease $S^{conf}(T_g)$ and in parallel increase the glass transition temperature, as shown by Le Losq et al. (2014) in the Na_2O - Al_2O_3 - SiO_2 system. A similar effect is potentially observed here, with the presence of $[5]Al$ possibly producing a decrease in $S^{conf}(T_g)$ and also an increase in T_g because $[5]Al$ and $[4]Si$ do not interact and do not mix randomly. One should be careful with extrapolating such behavior because it is only valid close to the glass transition. Indeed, at high temperatures (close or above the liquidus), Neuvill et al (2008b) proposed that $[5]Al$ plays a role similar to $[5]Si$, a transient species ensuring the exchange of O atoms during viscous flow according to Stebbins (1991). Similarly, in polymerized aluminosilicate networks, we can infer that five-fold coordinated network formers (Si or Al) ensure cationic mobility at high temperature. In depolymerized networks, Q^1 , Q^2 and Q^3 units ensure

network mobility (Stebbins, 1991; Farnan and Stebbins, 1990, 1994; Neuville et al., 2008b, Le Losq et al., 2014). Such assumptions agree well with the data near the C3A domain (see section 4). From Figure 21, it is further possible to make the assumption that $^{[5]}Al$ plays a different role in silica-rich or alumina-rich glasses and melts. Indeed, in silica-rich networks (>60 mol% SiO_2), $^{[5]}Al$ increases network connectivity and this results in a decrease in $S^{conf}(T_g)$. In an aluminate-rich network, a few percent of $^{[5]}Al$ plays a role similar to $^{[4]}Al$, which exchanges randomly with Si.

Other systems show behaviors along the tectosilicate join similar to that previously presented for $CaAl_2O_4-Si_2O_4$. In particular, Raman spectra along all tectosilicate joints in the LAS, NAS, KAS, MAS, SAS, BAS systems show similar variations. They all present a linear decrease in the frequency of the bands assigned to T-O symmetric and asymmetric stretching of Q^4 units with decreasing SiO_2 concentration. This correlates very well with increasing ^{27}Al NMR chemical shifts with decreasing silica content in tectosilicate glasses (see Figure 12 previous chapter, Drewitt et al., 2021). Those Raman and NMR variations arise from the variation of the T-O distance between Si-O and Al-O, as the Si-O distance is shorter than the Al-O distance.

In the Figure 22, the viscosities of tectosilicate melts with 75 and 50 mol % silica are plotted as a function of T_g/T . The strongest glass in the Angell plot corresponds to the SiO_2 composition, and the more fragile the composition, the greater the curvature. The fragility of melts decreases from Mg, Ca, Sr to Ba for the alkaline-earth elements. Alkali tectosilicate compositions show similar fragilities at 50 mol % silica (Figure 22B). At 75 mol % (Figure 22A), Li and K tectosilicate compositions also present a similar fragility while the Na-tectosilicate composition is the strongest one.

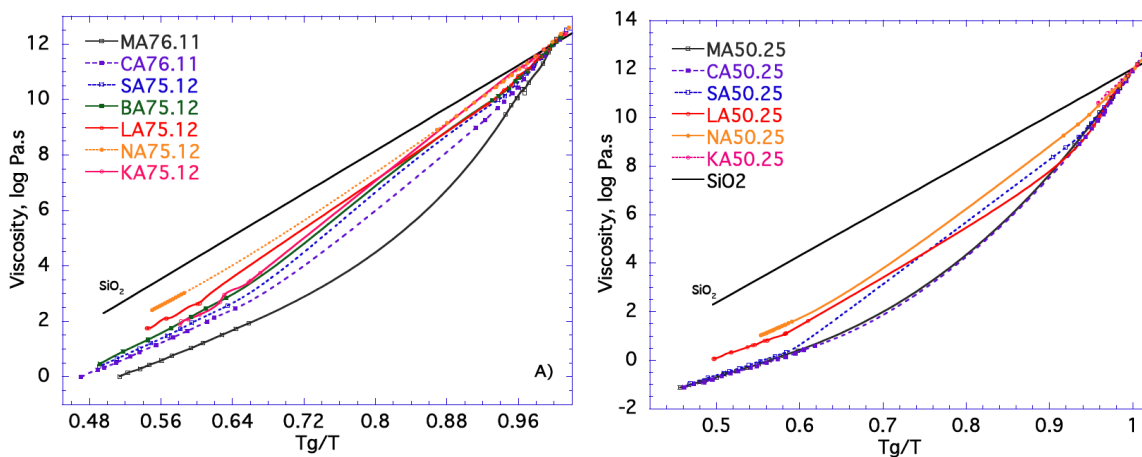


Figure 22: Viscosity of MAX.Y melts versus T_g/T , with $M = Li_2, Na_2, K_2, Mg, Ca, Sr, Ba$. $X = SiO_2$, $Y = Al_2O_3$, and $MO = 100 - (SiO_2 + Al_2O_3)$ in mole%, A) for 75 mole % of SiO_2 , and B for 50 mole % of

SiO_2 . (data are from [Le Losq, 2012](#) for LAS, [Le Losq et al., 2017](#), for NAS and KAS, [Neuvill 1991](#) for CAS and MAS, [Novikov et al.2017](#), for SAS, and [Novikov, 2018](#) for BAS).

3.5 From the CAS system to other chemical systems

For tectosilicate glasses and melts, the structure seems to present a random distribution and random substitution between SiO_4^{4-} and AlO_4^{5-} units, in violation of the Loewenstein's aluminum avoidance principle ([Loewenstein, 1954](#)). The Si/Al substitution is clearly confirmed by the linear variation of the chemical shift of the ^{27}Al NMR and of the frequency of the Raman band in the $800 - 1200\text{cm}^{-1}$ domain. We can propose that alkali or alkaline-earth elements are in the free-volumes of the $\text{Al}_x\text{-Si}_{(1-x)}\text{O}_2$ network, ensuring the charge compensation of AlO_4^{5-} units. The Figure 23 shows the proportion of $^{[5]}\text{Al}$ in tectosilicate glasses versus SiO_2 content. The proportion of $^{[5]}\text{Al}$ is a maximum in the middle of the ternary system, in the peraluminous domain for all ternary systems MAS (with $\text{S}=\text{SiO}_2$, $\text{A}=\text{Al}_2\text{O}_3$, $\text{M}=\text{Li}_2\text{O}$, Na_2O , K_2O , MgO , CaO , SrO , BaO), and increases with the decrease of z/r^2 ([Lee and Stebbins, 1999, 2000](#), [Allward et al, 2003, 2007](#), [Stebbins, 2008, 2016](#); [Stebbins and Farnan, 1992](#); [Stebbins et al., 1999](#); [Neuvill et al., 2004, 2006, 2008, 2010](#), [Le Losq et al., 2017](#), [Lee et al., 2016](#); [Novikov et al., 2017](#), [Stebbins et al., 2008](#), [Lee et al., 2009](#), [Hiet et al., 2009](#), [Alu et al., 2018](#)). In the NAS and KAS system, $^{[5]}\text{Al}$ is not detected in tectosilicate compositions, and it is only present in peraluminous domain ([Allwardt, et al., 2003](#); [Le Losq et al., 2014](#)).

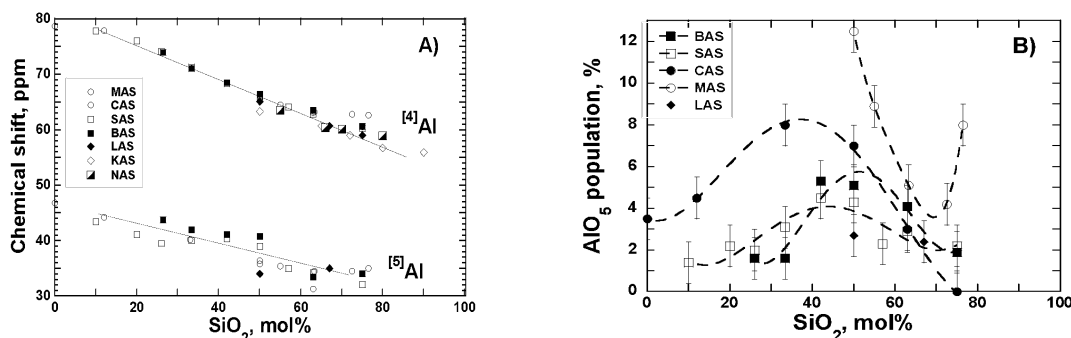


Figure 23: A) NMR chemical shift of $^{[4]}\text{Al}$ and $^{[5]}\text{Al}$ and B) proportion of $^{[5]}\text{Al}$ in LAS, MAS, CAS, SAS, BAS system along the join $R=1$, the tectosilicate join (note that there is not $^{[5]}\text{Al}$ in Na and K tectosilicate glasses). Data from [Le Losq \(2012\)](#), [Neuvill et al., \(2008, 2006\)](#), [Novikov et al., \(2017\)](#), [Novikov, \(2017\)](#). Error bar are smaller than symbol size.

In tectosilicate glasses, the NMR chemical shifts of $^{[5]}\text{Al}$ and $^{[4]}\text{Al}$ decrease linearly with the silica content ([Figure 23A](#)), due to an increase in the mean T-O-T inter-tetrahedral angle as Al substitutes for Si in tetrahedral units ([Stebbins and Farnan, 1992](#); [Neuvill et al., 2006, 2008](#)). At a given silica content,

small variations in the ^{27}Al NMR chemical shift are observed, and correlate with the electronegativity and size of the first-neighbor metal cations in the glass structure. These variations arise from small changes in the Si/Al distribution and in the mean inter-tetrahedral T-O-T angle as the metal cation field strength and property changes.

In peralkaline glasses, the degree of polymerization plays a small role on the ^{41}Al environment. Mysen et al. (2003) showed that, in sodium aluminosilicates, Al remains mostly in Q^4 units. Using Raman and viscosity data, Le Losq et al. (2014) corroborated such a view for this system. Other studies suggest that this remains true probably in most aluminosilicate compositions at silica contents higher than 40 mol% (Neuvill et al., 2004b, 2008, Novikov et al., 2017); a strong preference of ^{41}Al for the most polymerized structural units is expected according to experimental (Mysen et al., 1981; Cormier et al., 2000; 2005, Neuvill et al., 2004b, 2008, Novikov et al., 2017) and numerical studies (Cormier et al., 2003, Cormier 2019).

3.6 Alkaline-earth mixing in aluminosilicate glasses and melts

The substitution of one alkaline-earth element by another has little affect on glass Raman spectra (Merzbacher and White, 1991, Neuvill et al. 2008b). The Si/Al substitution actually has a greater influence on the aluminosilicate network than the alkaline-earth element substitution (see also chapter 3, this volume). In the case of Ca-Mg aluminosilicates, this agrees with an ideal mixing of Ca and Mg in aluminosilicates as inferred from macroscopic viscosity measurements (Neuvill and Richet, 1991). This random distribution was confirmed by ^{17}O NMR data, which show that NBOs, are equally associated with Mg or Ca cations (Farnan and Stebbins, 1990, 1994; Allwardt and Stebbins, 2004). However, the network is more perturbed by the presence of Mg than by Ca because of the higher ionic field strength of Mg (Navrotsky et al., 1982; Roy and Navrotsky, 1985) and the Ca/Mg coordination number should play an important role. The coordination number of Mg is well characterized (Kroeker and Stebbins, 2000, Trcera et al. 2009, Fiske and Stebbins, 1994): it is 4 in silicate glasses, and increases up to 5-6 in aluminosilicate glasses, in agreement with glass molar volumes variations (Figure 15). Turning to Ca, in silicate glasses it is mainly located in distorted sites with 6-7 oxygen neighbors (Cormier et al., 2003, Neuvill et al., 2004b, 2008, Cicconi et al., 2016). Its CN increases with substitution by Al as shown by the variation of the molar volume between CaSiO_3 and Al_2SiO_5 glasses (Figure 17). At high Al_2O_3 content ($R \geq 1$), in Ca/Mg aluminosilicate glasses, Ca is in 7-8 fold coordination (Shimoda et al., 2007a) whereas Mg is in 6-fold coordination (Shimoda et al., 2007b, Trcera et al. 2009), as corroborated by data like molar volume, X-ray diffraction spectra, XANES spectra at the Ca and Mg K-edge, and ^{43}Ca and ^{25}Mg NMR experiments (Kroeker and Stebbins, 2000, Trcera et al. 2009, Fiske and Stebbins, 1994).

The above discussion raises a paradox: Ca and Mg always have different coordination numbers but seem to mix randomly in silicate and aluminosilicate melts and glasses. First, this may highlight the difficulty determining alkali or alkaline-earth environments that are loosely held in large sites with fairly irregular coordination (Cormier and Neuville, 2004, Neuville et al., 2008). Nevertheless, there is a general consensus for lower coordination for Mg than Ca. The low Mg coordination number and the possibility to have MgO₄ tetrahedra in Mg-silicate glass suggest that Mg will not be available for charge compensation of Al in tetrahedral position. Alternatively, its small size favors its localization in network tetrahedral cavities, yielding important distortions of the aluminosilicate network to accommodate such Mg coordination, which could result in the formation of highly-coordinated Al. This interpretation might justify the modifier role played by Mg despite their low coordination (Shimoda et al., 2007b, Guignard and Cormier, 2008, Trcera et al. 2009), in agreement with recent NMR data suggesting that substitution of Na by Mg in aluminosilicates in glass promotes melt depolymerisation (Sreenivasan et al. 2020). Similar observations can be made regarding the role of Zn in silicate or aluminosilicate glasses (Novikov, 2017). The fact that the viscosities of the Mg or Zn glass compositions are lower than that of Ca glass compositions for the same amount of SiO₂, confirms the role of network modifier of Mg and Zn while they are preferably in 4-fold coordination. (Neuville, 1992; Neuville and Richet, 1991; Toplis and Dingwell, 2004, Neuville et al., 2008).

In mixed CMAS glasses, Raman, NMR and neutron diffraction data (Allwardt and Stebbins, 2004, Allwardt et al., 2003, Lee et al., 2003, Lee and Stebbins, 2003; Neuville et al., 2006, 2008, Guignard and Cormier, 2008; Cormier and Coello, 2013, Cormier 2019) are in good agreement with random mixing between Ca and Mg in the melt structure, as initially proposed by Neuville and Richet (1991). This implies that intermediate CMAS melts can be considered as being derived from the random mixing of their CAS and MAS end-members, e.g. Ca₃Al₂Si₃O₁₂ - Mg₃Al₂Si₃O₁₂ for garnet-like glasses and CaAlSi₂O₈ - MgAlSi₂O₈ for anorthite-like compositions. An interesting observation is that Ca/Mg mixing shows a random distribution and the network polymerization is not really affected by the Ca/Mg substitution, but the proportion of ⁵¹Al increases with the Ca/Mg substitution (Neuville et al., 2008b).

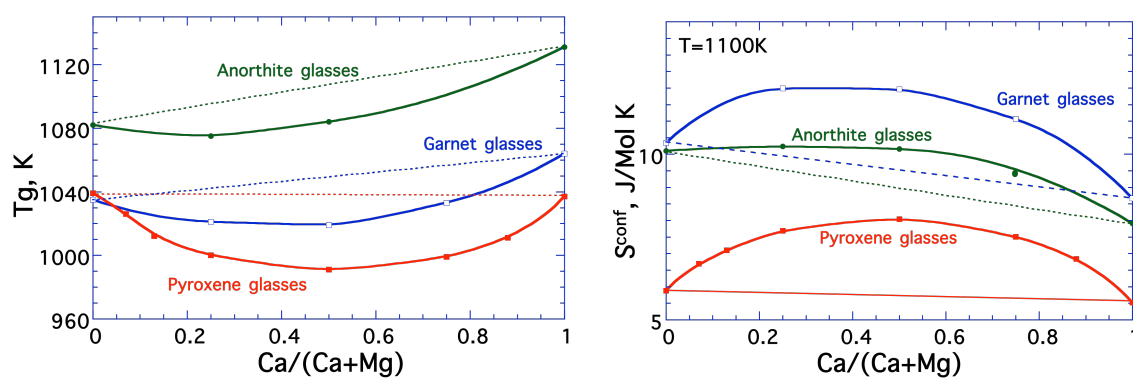


Figure 24: A) glass transition temperature and B) configurational entropy at 1100 K upon Ca/Mg mixing in pyroxene ($\text{CaSiO}_3\text{-MgSiO}_3$), garnet ($\text{Ca}_3\text{Al}_2\text{Si}_3\text{O}_{12}\text{-Mg}_3\text{Al}_2\text{Si}_3\text{O}_{12}$) and anorthite ($\text{CaAl}_2\text{Si}_2\text{O}_8\text{-MgAl}_2\text{Si}_2\text{O}_8$) glass compositions, with respectively 0, 12.3 and 25.0 mol% Al_2O_3 . Data for pyroxene and garnet glasses are from Neuville and Richet (1991) and original data for anorthite glasses.

Unfortunately, knowledge regarding the mixing between Sr/Ca, Ba/Mg or more generally between an alkali with an alkaline-earth element in aluminosilicate compositions remain scarce, and it is not possible to propose a general conclusion. It is possible to say that Si and Al are almost randomly distributed in tectosilicate glass compositions, and metal cations are located in cavities of the 3D polyhedral network and ensure the charge compensation of the AlO_4^{5-} tetrahedra. This assumption is in good agreement with the linear variation of the $T\text{-O}$ distances from 1.66 up to 1.76 Å between SiO_2 and CaAl_2O_4 glasses (Figure 13b, chapter 3, this issue). The $T\text{-O}$ distances are also well correlated with the linear variation of the NMR chemical shift of ^{27}Al and the linear variation of the frequency of the Q^4 species along the tectosilicate joint (compiled in Figure 13a,b chapter 3, this issue). They may mix randomly like Ca and Mg, but a non-random distribution with increasing network segregation has also been reported in alkali tectosilicate melts (Le Losq et al., 2017). This highlights the complexity of the structure of aluminosilicate melts and glasses, a problem when considering the modelling of their properties. Other problems are the presence of excess NBOs, $^{[5]}\text{Al}$ and tri-coordinated oxygens (Stebbins and Xu, 1997, Le Losq et al. 2014). For instance, the presence of $^{[5]}\text{Al}$ can be important, reaching concentrations up to 12% of the Al_2O_3 in the middle of the MAS ternary system, and the presence of $^{[5]}\text{Al}$ results in increasing the glass transition temperature and decreasing configurational entropy. Metal cations could be located in cavities of the 3D polyhedral network, in a manner consistent with the Compensated Continuous Random Network model (Greaves and Ngai, 1995; Greaves and Sen, 2007; Le Losq et al., 2017); they ensure the charge compensation of the AlO_4^{5-} tetrahedra. In the potassium aluminosilicate ternary, a non-random distribution of Si and Al between framework units may be accompanied by an increasing segregation of metal cations in clusters or channels. This has been proposed to explain the variations in

the structure and properties of potassium tectosilicate melts, for instance (Le Losq and Neuville, 2013; Le Losq et al., 2017). This is a problem when considering the modelling of their properties, and a large number of approximation may be necessary as we shall see in the later section.

3.7 Models for alkali aluminosilicate melts

The above sections allow us to have a glimpse into the complexity of aluminosilicate melts. Even leaving aside the potential effects of parameters like $^{[5]}\text{Al}$, there is no data for the distribution of Q^n species in aluminosilicate melts because ^{29}Si NMR spectroscopy loses resolution with aluminosilicate compositions due to Si-Al interactions resulting in significant signal broadening and increased complexity of the peaks. Mysen et al. (2003) quantified the fractions of the different Q^n species in Na aluminosilicate glasses by modelling the ^{29}Si NMR data. Their findings allow us to constrain the K_{eq} value for the dissociation $2 Q^3 = Q^2 + Q^4$ in sodium aluminosilicate glasses (Figure 25). We observe that Al introduction into the glass network promotes the dissociation of Q^3 units.

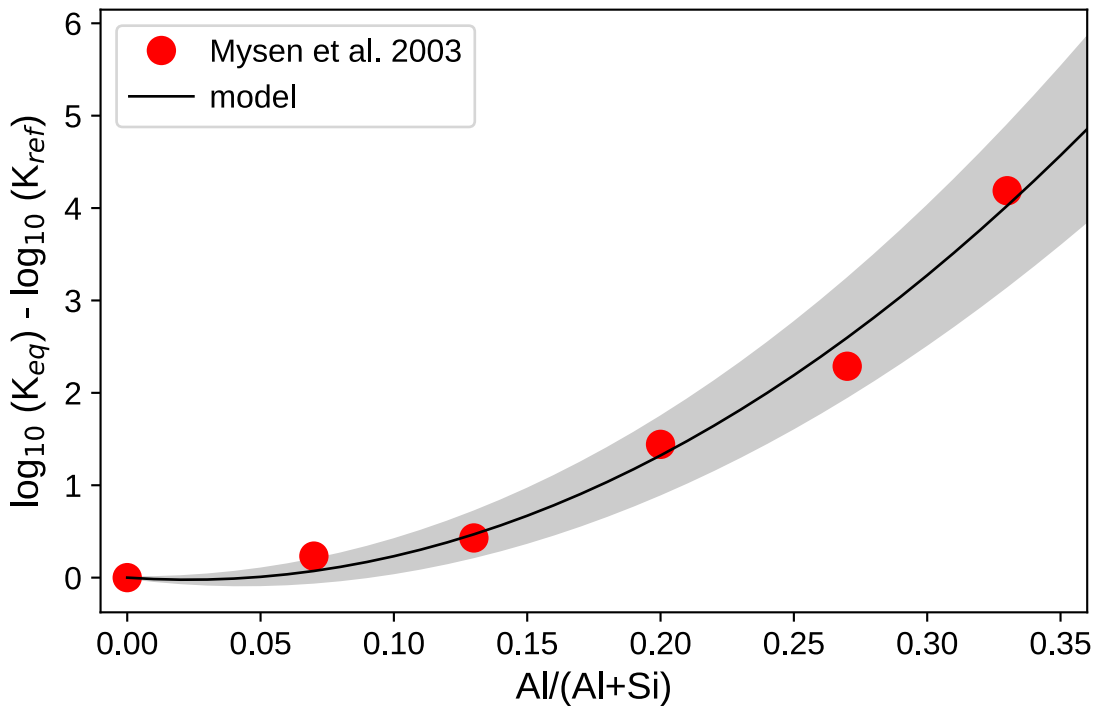


Figure 25: Relative value of the equilibrium constant K_{eq} of the reaction $2 Q^3 = Q^2 + Q^4$ obtained from Q^n data in Na aluminosilicate melts calculated by Mysen et al. (2003), reported as a function of the $\text{Al}/(\text{Al}+\text{Si})$ ratio of glasses. K_{ref} is the value of K_{eq} in the glass without Al. The black line represents a polynomial modelling of the data, and the grey areas are the two-sigma confidence interval of the model predictions. Using this model, it is possible to estimate the Q^n distribution in alkali aluminosilicate glasses, but one must note that such relation remains unknown for other elements like K, Ca or Mg.

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This result is the only experimental one to the knowledge of authors that could be used to model the Q^n distribution in aluminosilicates based on the principle of using partial K_{eq} values as highlighted in section 2.5. Other models (Doweidar, 1999 for example), like the associate solution one, are purely theoretical and remain unconstrained by experimental data, as they are not available. This does not prevent one from building models of glass and melt properties. For instance, Doweidar (1999) simply used a binary model for calculating the density of $\text{Na}_2\text{O}-\text{Al}_2\text{O}_3-\text{SiO}_2$ glasses from partial molar values of different $\text{Na}-Q^n$ units. The structural model employed there is clearly not appropriate, but interestingly, this does not prevent a good prediction of glass densities. A better attempt to link melt composition, structure and viscosity is that of Starodub et al. (2019), who calculated the Q^n distribution in $\text{Na}_2\text{O}-\text{K}_2\text{O}-\text{Al}_2\text{O}_3-\text{SiO}_2$ melts from an associate solution model, and then linked the fractions of Q^n units and the concentrations of cations to the parameters of the Avramov-Milchev equation for calculation of melt viscosity. This is an interesting model, but it suffers from two problems. First, they did not select the viscosity data used for fitting their model, as they used the large SciGlass database, and the results of their model will thus suffer from the propagation of large errors affecting some experimental data. Secondly, the Avramov-Milchev (Avramov and Milchev, 1988) equation works well but it is less attractive than the Adam-Gibbs theory because it does not allow one to link the chemical, structural, thermodynamic and dynamic dimension of the problem.

An extension of the model presented in section 2.5 could allow one to circumvent such caveats. We will explore if such a model is possible to implement below for the $\text{Na}_2\text{O}-\text{K}_2\text{O}-\text{Al}_2\text{O}_3-\text{SiO}_2$ system. A first step would be to combine the silicate structural model (section 2.5) to the relationship presented in Figure 25. This will allow one to approximate the Q^n speciation in aluminosilicate glasses. Then, we need to know the effects of Na-K mixing in aluminosilicate melts and of Al concentration on the glass configurational entropy. Combining viscosity and Raman data with thermodynamic calculations, Le Losq et al. (2014) have shown that Al introduction in NaSi_2O_7 yields a rapid network polymerisation (Figure 26A) that is accompanied by a jump in glass configurational entropy and an increase in glass T_g (Figure 26B). At $\text{Al}/(\text{Al}+\text{Na})$ higher than 0.5, we see a continuous decrease in $S^{conf}(T_g)$ that is accompanied by increasing T_g .

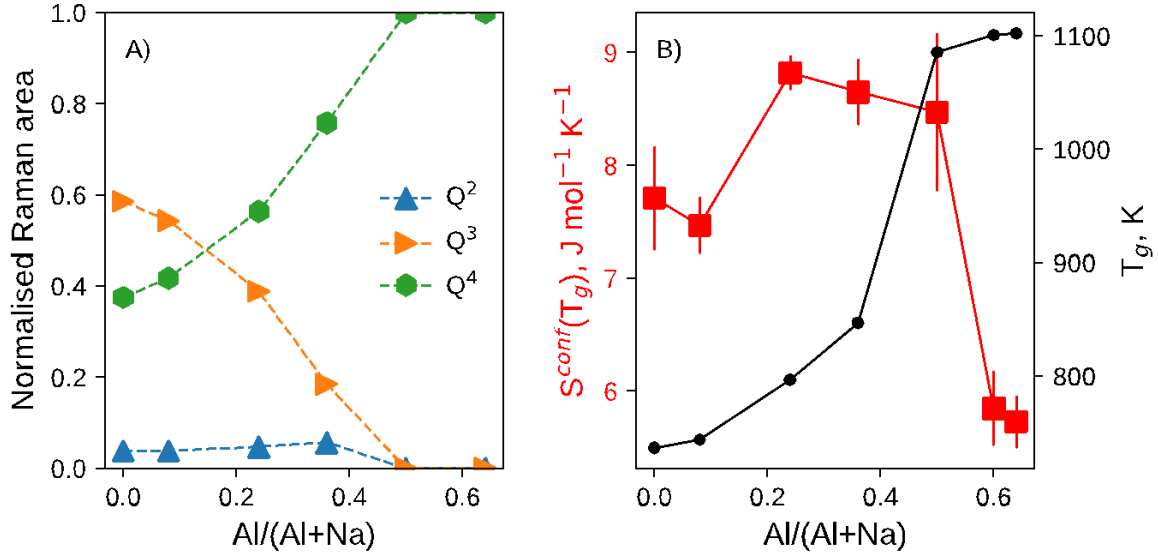


Figure 26: A) Fractions of the Q^n units Raman signal, and B) entropy and glass transition temperature of Na aluminosilicate glasses with 75 mol% SiO_2 but varying $\text{Al}/(\text{Al}+\text{Na})$ ratios. Data from Le Losq et al. (2014).

Variations in $S^{conf}(T_g)$ are thus complex functions of the melt $\text{Al}/(\text{Al}+\text{M})$ ratio. As an approximation, one could assume that Al and Si mix randomly in Q^4 units in the silicate network, as proposed for Na tectosilicate melts by Neuville and Mysen (1996). Note that it is important to remember that unlike the Ca or Mg aluminosilicate system, the ternary system $\text{Na}_2\text{O}-\text{Al}_2\text{O}_3-\text{SiO}_2$ does not contain Al in 5-fold coordination except in the peraluminous domain (Lee et al., 2003, Le Losq et al., 2014). This hypothesis implies that the Lowenstein Al-O-Al avoidance rule will not be respected, and is thus in contradiction with the interpretation of NMR data of Na aluminosilicate glasses (Lee and Stebbins, 1999, Lee et al., 2003). However, this hypothesis allows reproduction of the variations in configurational entropy observed when Al substitutes Si in Na aluminosilicate melts (Neuville and Mysen, 1996; Le Losq, 2012). One may thus consider it as adequate for the aim of building a model of configurational entropy variations. An additional point is to consider that Q^4_{Al} charged-balanced by Na ($Q^4_{\text{Al,Naenv}}$) or K ($Q^4_{\text{Al,Kenv}}$) probably present different partial configurational entropies (respectively $S^{conf}_{Q^4_{\text{Al,Naenv}}}$ and $S^{conf}_{Q^4_{\text{Al,Kenv}}}$). S^{topo} (eq. 16) could then be written as:

$$S^{topo} = \sum_{n=2}^3 x_{Q^n_{\text{Si,Naenv}}} S^{conf}_{Q^n_{\text{Si,Naenv}}} + \sum_{n=2}^3 x_{Q^n_{\text{Si,Kenv}}} S^{conf}_{Q^n_{\text{Si,Kenv}}} + x_{Q^4_{\text{Si}}} S^{conf}_{Q^4_{\text{Si}}} + x_{Q^4_{\text{Al,Naenv}}} S^{conf}_{Q^4_{\text{Al,Naenv}}} + x_{Q^4_{\text{Al,Kenv}}} S^{conf}_{Q^4_{\text{Al,Kenv}}} \quad (32)$$

From this equation, we need to discriminate the fraction of charge-compensator and network modifiers (Na, K), in order to calculate the variation in the fractions of the Q^n units. This could be done easily assuming that the fractions of charge-compensator and network modifiers Na and K are just equal to the weighted average fractions of Na and K. This assumes no preference of Na and K to specific environments, an hypothesis that surely could be challenged.

We see that many hypotheses have been already proposed in order to start building a model to express $S^{conf}(T_g)$. This is without even considering the calculation of any entropy resulting from cationic mixing, S^{mix} . The picture becomes even more complex at this point. Indeed, while variations of $S^{conf}(T_g)$ upon mixing Na and K in silicate melts can be explained by a random mixing of those cations (Richet, 1984; see section 2.4, eq. 18): mixing Na and K in the presence of Al results in a different behaviour. Entropy variations upon mixing Na and K in tectosilicate melts ($R=1$) are far from ideal (Le Losq and Neuville, 2013). Viscosity, Raman, NMR and molecular dynamic simulation data all indicate that Na and K actually are present in different environments in tectosilicate melts (Le Losq et al., 2017), with K tending to be present in large percolation channels in the polyhedral Al-Si network. This was recently challenged by a new analysis of viscosity data along the $\text{SiO}_2\text{-(Na,K)AlO}_2$ binary by Robert et al. (2019). Those authors fitted viscosity data of mixed Na-K tectosilicate melts, leaving the A_e parameter (eq. 6) free. This yields results where it is possible to fit viscosity upon Na-K mixing in tectosilicate melts assuming a random Na-K distribution, like in silicate melts. In such cases, most of the variations in viscosity seemed to be explained by variations in A_e . While this is an interesting take on this problem, this is probably not a good representation of the reality. Indeed, many authors demonstrated that the pre-exponential term in viscosity equations converge to a singular value (Persikov, 1991; Giordano and Dingwell, 2003, Giordano et al., 2008, Giordano and Russell 2018). The A_e parameter is unlikely to vary largely in aluminosilicate melts (e.g. Russell and Giordano, 2005, 2017 and references therein). It thus seems improbable that A_e varies largely upon substitution of Na by K at fixed SiO_2 and Al_2O_3 contents. Raman spectroscopy data and molecular dynamic simulations (Le Losq and Neuville, 2013; Le Losq et al., 2017, Le Losq et al., 2019) further indicate that Na tectosilicate and K tectosilicate melts do not have the same structure, such that Na and K do not mix randomly in such systems. The mixing of Na- and K-tectosilicates thus appears to be of a mechanical mixing, and should accordingly yield a pseudo-linear variation of $S^{conf}(T_g)$. This agrees with conclusions from Le Losq and Neuville (2013) as well as Le Losq et al. (2017), which report pseudo-linear variations in $S^{conf}(T_g)$ upon Na-K mixing.

Considering that mixing Na and K compensators does not affect entropy, we need to consider three sources for the calculation of S^{mix} : (i) the ideal mixing of Si between different Q^n units, (ii) the

ideal mixing of Si and Al in Q^4 units and (iii) the ideal mixing of Na and K network modifiers. We thus have:

$$S_{Si-Al}^{mix} = -2 \frac{x_{Si \in Q^4} + x_{Al}}{x_o} R \left[\frac{Al}{Al+Si} \log \frac{Al}{Al+Si} + \left(1 - \frac{Al}{Al+Si}\right) \log \left(1 - \frac{Al}{Al+Si}\right) \right], \quad (33)$$

with $x_{Si \in Q^4}$ the mole fraction of Si in Q^4 units,

$$S_{Si}^{mix} = \frac{-x_{Si}}{x_o} * 2 * R * \sum_{n=2}^4 x_{Q_{Si}^n} \log(x_{Q_{Si}^n}), \quad (34)$$

and assuming a distribution of Na and K as charge-compensator and network modifier that varies linearly with Al/(Al+M), we finally have

$$S_{Na-Kmodifiers}^{mix} = MODS \times \left[-2R \frac{x_{Na} + x_K}{x_o} (X_K \log(X_K) + (1 - X_K) \log(1 - X_K)) \right], \quad (35)$$

with MODS the fraction of modifiers contributing to S^{mix} ; MODS is equal to $(1 - Al/(Na+K))$ for $Al < (Na+K)$, and is equal to 0 for $Al \geq (Na+K)$.

The sum of equations 32 to 35 allows the calculation of $S^{conf}(T_g)$. B_e can be calculated following a similar logic to that used for silicate melts (section 2.5):

$$B_e = \sum_{n=2}^3 x_{Q_{Si,Naenv}^n} B_{eQ_{Si,Naenv}^n} + \sum_{n=2}^3 x_{Q_{Si,Kenv}^n} B_{eQ_{Si,Kenv}^n} + x_{Q_{Si}^4} B_{eQ_{Si}^4} + x_{Q_{Al,Naenv}^4} B_{eQ_{Al,Naenv}^4} + x_{Q_{Al,Kenv}^4} B_{eQ_{Al,Kenv}^4} + K_1 S_{Si}^{mix} + K_2 S_{Na-Kmodifiers}^{mix} + K_3 S_{Si-Al}^{mix}, \quad (36)$$

with K_1 , K_2 and K_3 being proportionality constants accounting for the influence of the intrinsic entropy in B_e (section 2.4; Le Losq and Neuville, 2017). A_e is left as a common parameter.

Using the equations described above, and selecting viscosity data from the literature for parameter tuning via least-square (see Figure 27), we are able to build a model for the viscosity of melts in the $Na_2O-K_2O-SiO_2-Al_2O_3$ system, with compositions from 50 mol% (in aluminosilicates; 60 in silicates) to 100 mol% SiO_2 , and NBO/T ranging from 0 to 1.33. Root-mean-square-errors between viscosity predictions and measurements are equal or lower than 0.5 log Pa.s. This demonstrates that building a structural model of the properties of aluminosilicate melts is possible. However, the above model is

incomplete because it relies on many hypotheses, and does not take into account (yet) many specifics like the role of $^{[5]}Al$ on $S^{conf}(T_g)$. Despite this, this is an encouraging step toward the construction of structural models that usually achieve a high degree of precision in the predictions they make. New developments in modelling the melt structure and in linking it with melt properties will surely allow improvements of such structural models in the future.

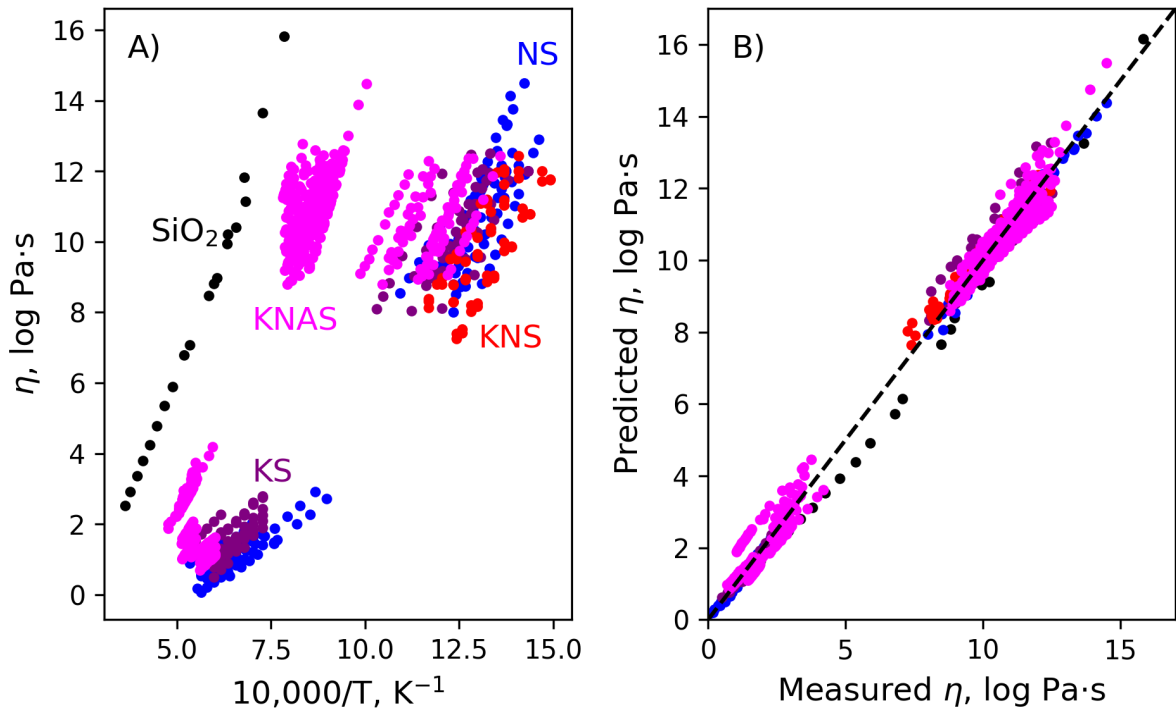


Figure 27. A) Viscosity data used for the model calibration. Data ($n=817$) are those compiled by Le Losq and Neuville (2017) for silicates (NS, KS, KNS and SiO₂), and for aluminosilicates (KNAS) they were compiled from Riebling (1966), Taylor and Rindome (1970), Urbain et al. (1982), Le Losq (2012), Le Losq and Neuville (2013), Le Losq et al. (2014), and Le Losq et al. (2017). The model parameters were optimised by No-U-Turn Markov Chain Monte Carlo sampling via the PyMC3 Python library. The root-mean-squared error between predictions and measurements is around 0.36 log unit.

4. Aluminate glasses and melts

Aluminate glasses and melts are of interest from a fundamental structural point of view: they contain only Al as a network former and, hence, allow better understanding of the role of Al in glasses and melts. The study of aluminate materials is also very important to understand the formation and evolution of parental bodies in the universe (Leger et al., 2009; Batalha et al., 2011). Aluminate materials further represent a technological material of interest because of their good IR transmission and ultralow optical losses (Higby et al., 1990) that make them attractive candidates for low loss optical fibers such as infrared

waveguides (Lines et al., 1989, King and Shelby, 1996). In addition, calcium aluminate glasses have been found to be photosensitive to ultraviolet radiations, leading to potential applications in photometric devices for information storage purposes (Hosono et al., 1985). They also have excellent mechanical properties, such that calcium aluminate glass fibres have been proposed for the reinforcement of cement composites (Wallenberger et al., 2004, El Hayek, 2017).

Looking at aluminate glasses allows one to better understand the fundamental role of aluminum in glasses and liquids, a task somehow difficult in aluminosilicate compositions because of the superposition of many complexities (see section 3). Numerous studies have probed the structure of calcium aluminate crystals, glasses and melts (e.g. McMillan and Piriou, 1983, Hannon and Parker, 2000, Benmore et al., 2003; Neuville et al., 2008, 2010, Licheron et al., 2011, Drewitt et al., 2012, 2017). From those, it appears that four compositional “domains” can be distinguished: Al_2O_3 , Al_2O_3 -CA, CA-C3A and C3A-CaO with $\text{CA}=\text{CaAl}_2\text{O}_4$, $\text{C3A}=\text{Ca}_3\text{Al}_2\text{O}_6$.

4.1 Al_2O_3

Crystalline Al_2O_3 has corundum structure with Al in 6-fold coordination (Rankin, 1915). Several studies, using ^{27}Al NMR at high temperature show that Al_2O_3 liquid is a mixture of 60% $^{[6]}\text{Al}$ - 40% $^{[4]}\text{Al}$. This was confirmed more recently by XANES at the Al *K*-edge and FDMNES simulation (Neuville et al., 2009). Recently, Shi et al. (2019) proposed, from new X-ray diffraction at high temperature measurements, that Al is essentially in five-fold coordination in the Al_2O_3 liquid; this interpretation is based on the Al-O distance. The ^{27}Al NMR chemical shift is 53 ppm (Couture et al., 1990; Florian et al., 1995, Ansell et al., 1997), a value that can be interpreted as $^{[5]}\text{Al}$, or as a mixture of $^{[4]}\text{Al}$ and $^{[6]}\text{Al}$. When assuming a mixture of $^{[4]}\text{Al}$ and $^{[6]}\text{Al}$, the high temperature ^{27}Al NMR data allow determining an Al NMR relaxation time in excellent agreement with the relaxation time obtained from viscosity measurements (Urbain, et al. 1982). Besides, when comparing XANES data at the Al *K*-edge measured at high temperature in the liquid state to spectra simulated using the FDMNES code (Joly, 2001), the assumption of a mixture of $^{[4]}\text{Al}$ and $^{[6]}\text{Al}$ provides the best reproduction of the experimental data (Neuville et al., 2009). From all the high temperature NMR and XANES spectroscopies and/or neutron diffraction measurements, it is at the moment preferable to consider that, in liquid alumina, Al is present as a mixture of $^{[4]}\text{Al}$ and $^{[6]}\text{Al}$.

We have essentially mentioned Al_2O_3 in its crystalline or liquid forms, however, some studies have shown that it is possible to obtain Al_2O_3 in its amorphous form, it is usually obtained by fast quenching and in thin film form (Lee et al., 2010; Kim et al., 2014). In that case, Al is essentially a mix of 4- and 5-fold coordination with a small amount of 6-fold coordination (Lee et al., 2010; Kim et al., 2014). The

difference between the Al coordination between glass and liquid state can be easily explained by the high cooling rate.

4.2 Al_2O_3 - $CaAl_2O_4$ compositions.

With the introduction of calcium (or others alkaline-earth elements) in alumina liquid, a first crystalline phase can be found: grossite, $CaAl_4O_7$, called CA2, with AlO_4 tetrahedra only in Q^4 units although one of the O atoms forms an “oxygen tricluster” which links three AlO_4 tetrahedra (Stebbins et al, 2001, Iuga et al., 2005, Neuville et al., 2010). With increasing Ca content up to the $CaAl_2O_4$ crystal composition, the amount of tricluster oxygen decreases and Al stays in four fold coordination. The $CaAl_4O_7$ composition always crystallizes while $CaAl_2O_4$ (CA) can be quenched to a glass by fast quench methods. In the CA glass, $[^4Al]$ is dominant and $[^5Al]$ representing 3,5 % of total Al (Neuville et al., 2006). In the CA liquid, $[^4Al]$ remains dominant but the proportion of $[^5Al]$ increases up to around 40 mol % near 2000 K, as observed from XANES at the Al K-edge and neutron diffraction (Neuville et al., 2008a, Drewitt et al., 2012). Oxygen triclusters also probably are present but remain unquantified during these changes. The coordination of calcium, which is 7-fold in the glasses, increases in the melts as indicated by the evolution of the pre-edge of Ca K-edge XANES spectra between room temperature and superliquidus temperatures (Neuville et al., 2008a).

4.3 CA-C3A compositions.

For compositions between $50 > Al_2O_3 > 25$ mol%, Al is only in tetrahedral coordination in both the glasses and the crystals, with a number of bridging oxygens varying between 4 to 2 respectively for CA and C3A (McMillan and Piriou, 1983, Neuville et al., 2004, 2006, 2010, Licheron et al. 2011, Drewitt et al., 2017). In C3A crystal or glass ($Ca_3Al_2O_6$), Al can be found in Q^2 units, and, with increasing temperature, Al K-edge XANES show very small changes (Neuville et al., 2008a). The later observation implies that Al stays in Q^2 species with increasing temperature without creation of $[^5Al]$. The cationic mobility is only insured by Q^2 species, and not by highly coordinated units as proposed by Stebbins and Farnan (1992) for polymerized melts. Calcium cations in crystalline C3A and glass is in 6-fold coordination, and, near 2000 K, a small pre-edge in Ca K-edge XANES spectra is observed; this implies that the Ca coordination increases with increasing temperature (Neuville et al., 2008a).

4.4 C3A-CaO compositions

For compositions with $Al_2O_3 < 25$ mol%, no glass or stable crystalline phases are observed. The knowledge of Al and Ca coordination in such conditions remains very scarce, if not non-existent.

4.5 Link with observations in other binary systems

In $\text{SrO-Al}_2\text{O}_3$ and $\text{BaO-Al}_2\text{O}_3$ systems, the liquidus temperatures are higher than in the $\text{CaO-Al}_2\text{O}_3$ system. This observation seems logical since there is agreement with an increase of liquidus temperature with increasing the size of the alkaline-earth elements in alkaline-earth aluminosilicate systems (Médéric et al., 2004, Licheron et al., 2011, Shan et al., 2018). Turning to $\text{MgO-Al}_2\text{O}_3$, Mg^{2+} forms a very stable spinel with alumina; it is not possible to quench the MgAl_2O_4 composition or other magnesium aluminate composition even with using fast-quench methods. In MgAl_2O_4 crystals, Mg and Al are in 4 and 6-fold coordinations, respectively, at room temperature and that can change with temperature (Andreozzi et al., 2000, Neuville et al., 2009).

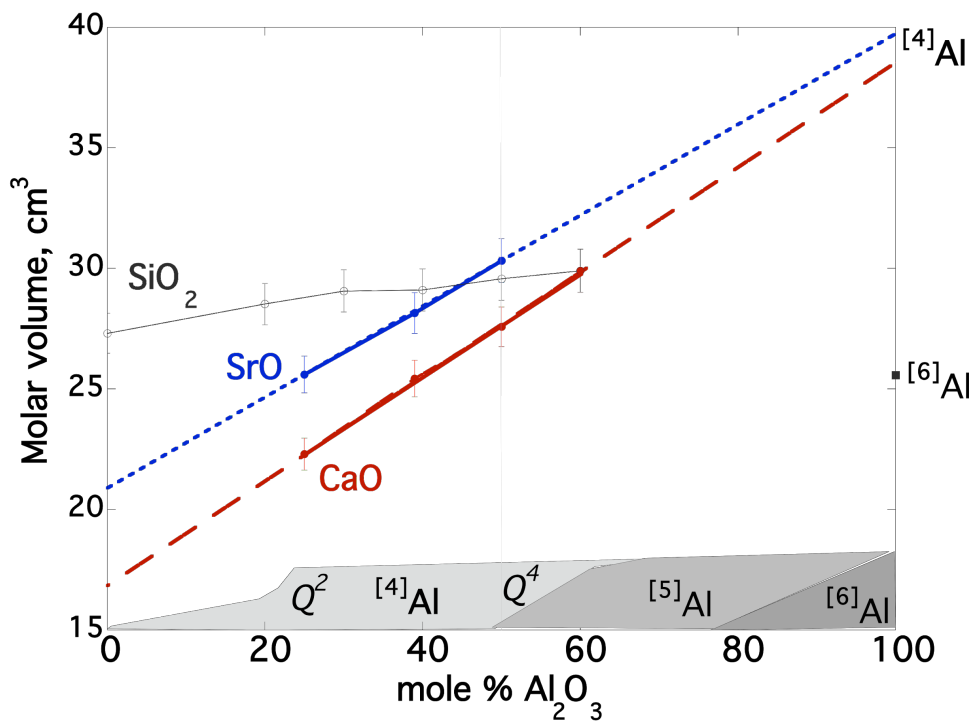


Figure 28: Coordination of Al between $\text{MO-Al}_2\text{O}_3$ and variation of molar volume of aluminate glasses. Molar volumes are calculated from the density measurements for $\text{Al}_2\text{O}_3\text{-SiO}_2$ (Wang et al., 2020), $\text{Al}_2\text{O}_3\text{-CaO}$ (Neuville et al., 2006, 2010, Licheron et al., 2011), and $\text{Al}_2\text{O}_3\text{-SrO}$ (Licheron et al., 2011). Solid lines correspond to fits of the data with a line, and dashed lines show the extrapolation of those fits to end members.

Figure 28 summarizes the evolution of the Al coordination in aluminate glasses and also shows how the molar volume of glass changes with Al_2O_3 content. The glasses CxAl or SxAl , with x varying between 3 and 1, are made by levitation and fast quench, and all these glasses show Al in 4-fold coordination (Neuville et al., 2006, 2008b, 2010, Licheron et al., 2010). For Ca or Sr aluminates, a linear variation of

the molar volume with the alumina content is observed between 25 and 50 mole percent of alumina, and Ca is in 6-fold coordination and Al in 4-fold coordination (Neuville et al., 2008b). If we consider linear extrapolations of these trends, it is possible to propose, considering only Al in four-fold coordination and Ca and Sr in 6-fold coordinations, that the respective partial molar volumes of Al_2O_3 , CaO and SrO are respectively of 39.1, 17.0 and 21.0 $\text{cm}^3 \text{mol}^{-1}$. Such estimations are in very good agreement with the values proposed by Robie (1978) for the molar volume of CaO and SrO.

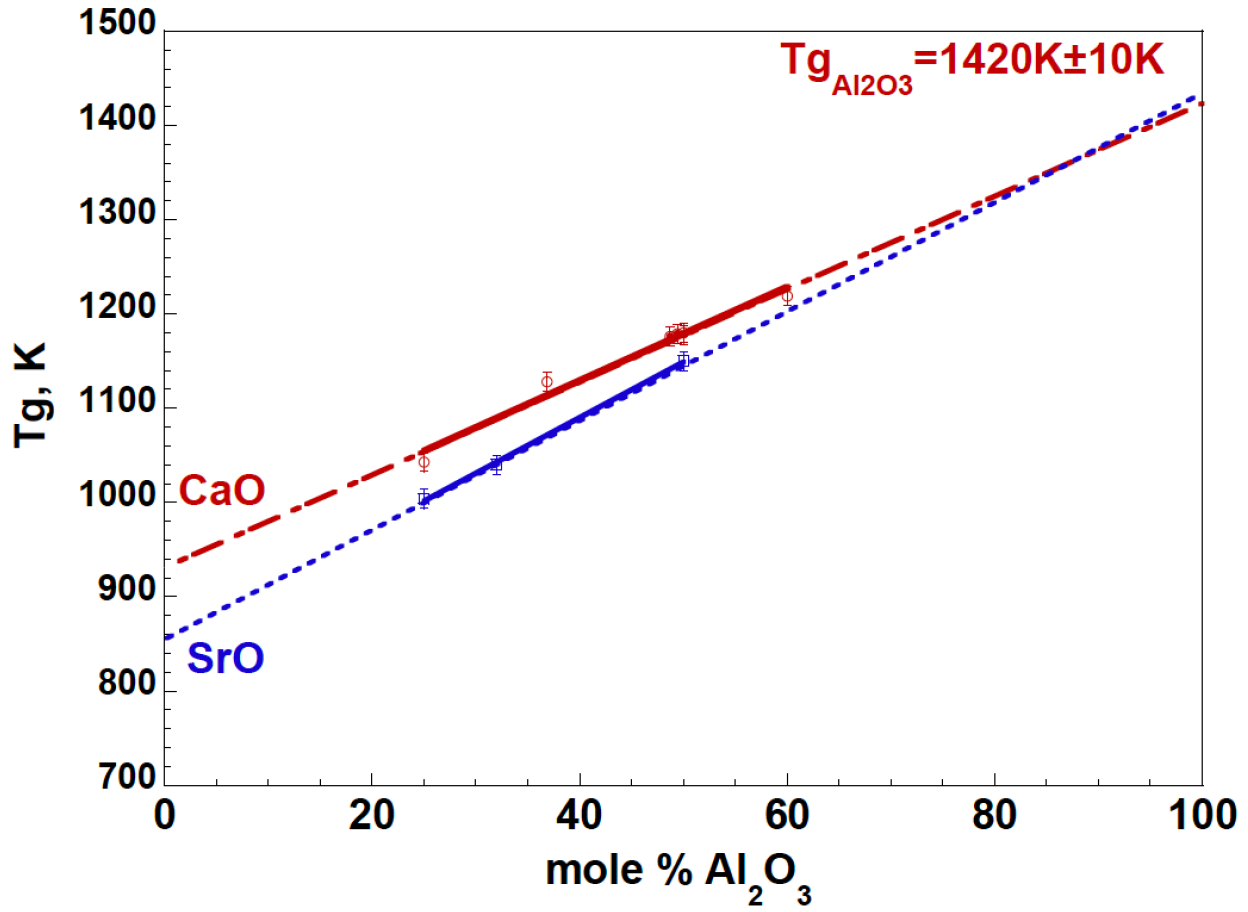


Figure 29: Glass transition temperature obtained from DSC measurements with and heating rate of 10°/min for aluminate glasses. Full line corresponds to linear variation fitting datas and dashed line corresponds to extrapolation up to end members.

To complete the observations made on the molar volume, the glass transition temperatures for CxA and SxA glasses were measured and plotted in Figure 28. The glass transition temperature of the different CxA and SxA glasses were obtained by DSC measurements with a heating rate of 10 °/min. Two linear trends are observed along the binaries CaO- Al_2O_3 and SrO- Al_2O_3 . They converge to similar values when extrapolated to the alumina end-member. From this, we can propose an “apparent” partial glass transition temperature for Al_2O_3 at 1420 ± 10 K. Following this logic, we also propose partial T_g s for CaO and SrO

in 6-fold coordination at 932 ± 10 and 857 ± 10 K. The estimated T_{gi} for Al_2O_3 with Al in four-fold coordination is in very good agreement with calculation of phase diagrams using CALPAHD program (Pitsch, et al. 2021 submitted).

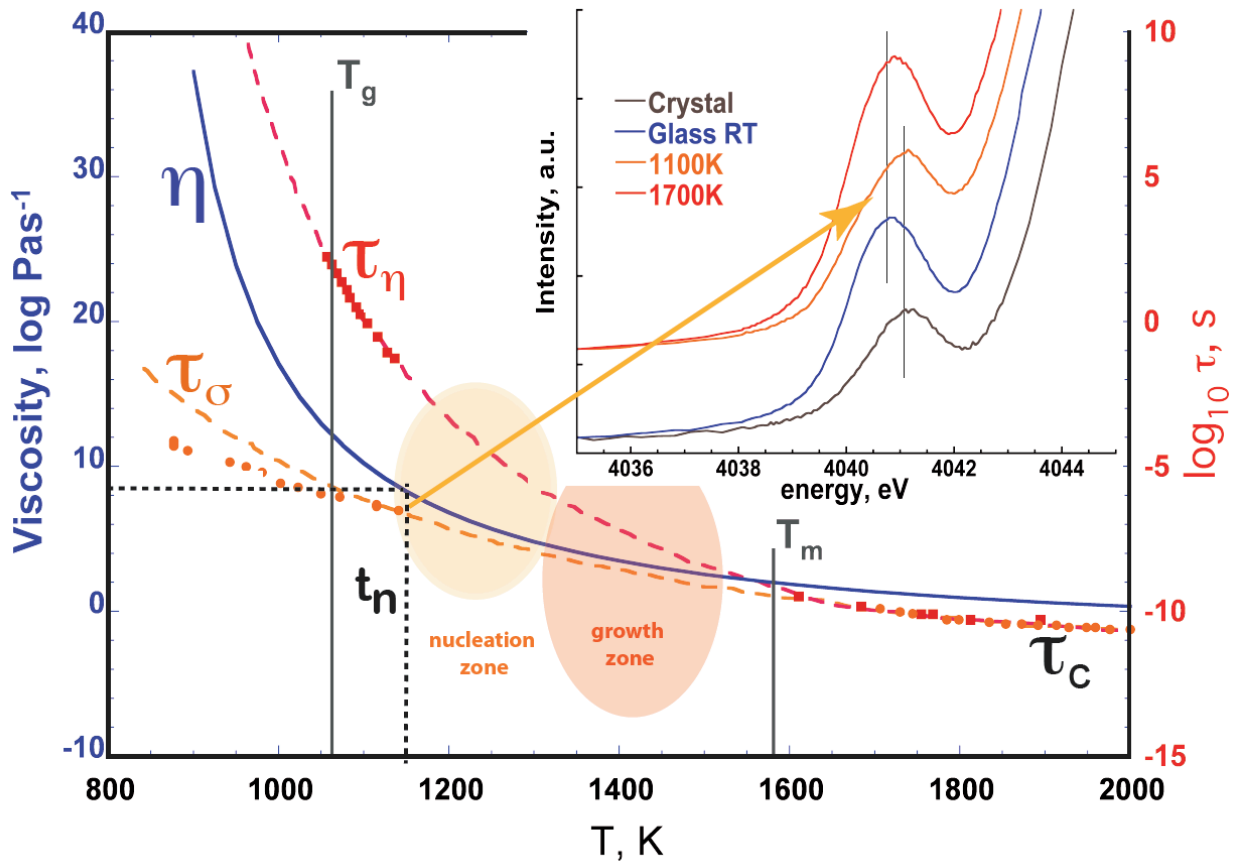
Recently Shan et al (2018) made some calcium-strontium peraluminous glasses at 58% mole of Al_2O_3 using levitation device and fast cooling rate. They investigated macroscopic properties of these glasses and they showed that an ideal mixing term between Ca and Sr clearly exist to explain Vickers micro-hardness and glass transition temperatures variation between Ca/Sr aluminate glasses.

5. Conclusion and perspectives

Since the 1980s, it has been possible to study the structure of a liquid at high temperature by several spectroscopic methods: Raman, NMR, X-ray absorption and X-ray and neutron diffraction techniques. This allows us to understand the local order and the medium range order directly in the liquid state and thus be able to link the high temperature structure to the measurements of properties directly accessible at high temperature, like heat capacity, viscosity, conductivity, diffusivity. We have seen that this kind of approach is very promising because it allows us to have a global view of a glass and a liquid and thus to better understand the structural changes that take place in a liquid. This improved understanding of silicate glasses and liquids has great applications in both volcanology and geochemistry where we are better able to understand changes in eruptive dynamism, mass transfers and also chemical balances during crystallization processes for example. But this work allows also a better understanding of industrial processes during melting, glass-forming or even during the processes of glass-ceramic.

To summarize the changes that can occur in a cooling liquid, Figure 30, obtained for a composition of $Ca_3Al_2Si_3O_{12}$ liquid, shows both the changes that occur in the liquid state and during its cooling. We can see that at high temperature, beyond the liquidus, T_m , 1580K, the relaxation times obtained from the ^{27}Al NMR, τ_{NMR} , measurement of conductivity, τ_σ , or viscosities, τ_η converge and are similar. It simply means that in the liquid state all cations whatever their roles in the melts, move in relation to each other. From structural point of view, it is a liquid state, characterized by the particularly intense pre-edge of the XANES spectra at the Ca K -edge at high temperature. On the other hand, when the temperature decreases, there is a decoupling between the various relaxations, this means that the atoms begin to move differently, the network modifiers like Ca, continue to move as shown by the relaxation τ_σ , while the relaxation calculated from the viscosity measurements, τ_η , shows a strong increase, meaning that the tetrahedral network composed of tetrahedral SiO_4 and AlO_4 moves more slowly, which translates into an increase in the viscosity curve. This decoupling between the various families of atoms appears at about 1100K for this chemical composition, several orders of magnitude

848 separate the relaxation times obtained from the conductivity measurements, τ_σ , from those obtained from
849 viscosity measurements, τ_η . These differences show, on the one hand, that the aluminosilicate network
850 freezes with a significant increase of τ_η , and viscosity and that calcium atoms put themselves in a position
851 close to that which they will have in the crystalline phase, as shown by the changes observed in the pre-
852 edge of the XANES spectra at the Ca K-edge. In the end, we find that close to T_g , the Ca is probably in a
853 crystallographic site almost identical to the one it will have in a crystal.
854



855
856 *Figure 30: viscosity of $\text{Ca}_3\text{Al}_2\text{Si}_3\text{O}_{12}$ liquid from Neuville and Richet (1991), relaxation τ_η is calculated*
857 *from viscosity measurements, τ_σ and τ_{NMR} , are from Gruener et al., (2001) and pre-edge XANES spectra*
858 *at the Ca K-edge are from Neuville et al., (2008b)*
859

860 Figure 30 completes the Figure 4 of this chapter, which helped to understand the links between potential
861 movements in a liquid, various entropic states and viscosity, but this latter figure indicates that at high
862 temperature all atoms can move quickly relative to each other and that at lower temperatures, atoms are
863 in different potentials energy states and that changes will take more time.

864 In 2021, it is therefore possible and necessary to understand the structure and properties of a
865 glass or/and a liquid throughout its history during a natural or industrial process. The current

computational techniques discussed in the following chapters can show and predict these different states, and it is easy to imagine that the new techniques based on Machine Learning will allow modeling of the structure, and physical properties and be able to predict them globally (Le Losq et al., 2021), this will have great consequences for both earth sciences and materials sciences.

6. References

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