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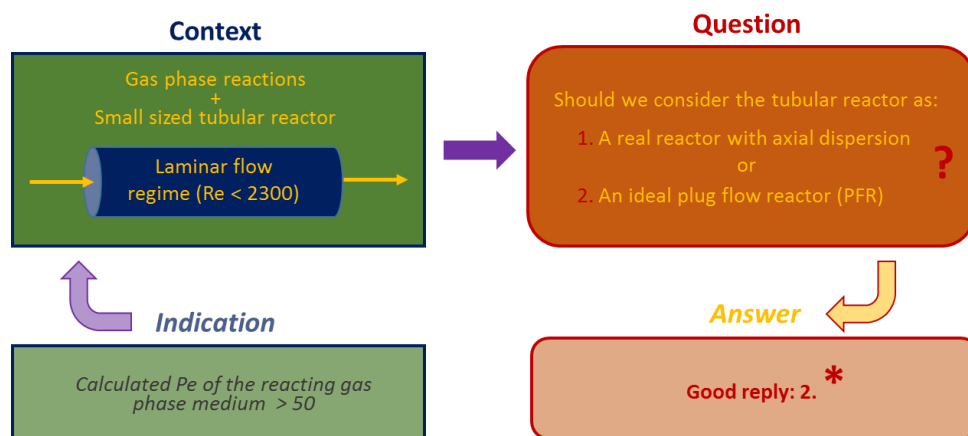
Julien Gornay, Edouard Plasari, Jean-François Portha*, Pierre-Alexandre Glaude, Francis Billaud, and Lucie Coniglio*

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Graphical abstract



* For details: Please see the manuscript.

Highlights

- Small sized tubular reactor appropriately designed to produce gas phase kinetic data
- Reason: axial dispersion coefficient much smaller in gas phase than in liquid phase
- Helpful tool for research dedicated to thermal conversion of liquid biomass

Design of a bench-scale tubular reactor similar to plug flow reactor for gas-phase kinetic data generation - Illustration with the octanoic acid pyrolysis

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Abstract

The material described in this article deals with waste conversion into energy vectors by pyrolysis, steam cracking or oxidation of liquid biomass, carried out at small to medium scale. The design of a bench-scale experimental setup devoted to gas phase kinetic data generation in a tubular reactor under laminar regime close to plug flow is detailed on the basis of a very simple approach. Validation of the designed bench-scale setup was successfully carried out within the context of octanoic acid pyrolysis by generating kinetic data with satisfactory measurement repeatability and material balances. The key of this positive result is that axial dispersion coefficient is much smaller in gas-phase than in liquid-phase, allowing thus the designed small sized tubular reactor to be close to the plug flow reactor. Such a feature of the axial dispersion coefficient is not well known by the wider public. Besides, octanoic acid was selected as surrogate for carboxylic acids because of their key role in various industrial applications (combustion of ethyl biodiesel; production of biofuel and biosourced chemicals).

Keywords: Octanoic acid; Pyrolysis; Kinetic data; Bench-scale setup; Design.

1. Introduction

Validation of models developed to depict accurately reacting systems is usually achieved when good matching between experimental information and simulated information is obtained [1]. However, this step cannot be conducted successfully if the experimental devices used to generate experimental information do not meet at least the theoretical hypothesis underlying the developed models.

Therefore, the purpose of this article is to describe step by step the design of a new bench-scale experimental setup devoted to generate kinetic information at constant near-atmospheric pressure. For the reasons mentioned previously, the design of the experimental setup should satisfy the assumptions of the models implemented in the most commonly used software for reacting system simulation, i.e. CHEMKIN II [2] or Aspen plus®, combined with EXGAS for the automatic generation of the detailed kinetic mechanisms relating to the reacting systems considered [3-7]. Regarding EXGAS, the reacting systems should involve reactions occurring in a homogeneous gas-phase medium at low to moderate pressures. Hence, no wall effects that would induce heterogeneous gas-solid catalytic or non-catalytic reactions between the inner surface of the reactor and the gas phase components should occur since these effects would be ignored by the model [8]. Regarding CHEMKIN II [2] or Aspen plus®, the implemented models are dedicated to representing the ideal flow patterns occurring either in the continuous stirred tank reactor (CSTR) or in the plug flow reactor (PFR). With the objective of a successful transfer to industrial reactors commonly found in pyrolysis, steam cracking or steam reforming processes, the option of a continuous flow tubular reactor was selected. This option goes of course with the requirement that flow is close to plug, despite the laminar regime imposed by a bench-scale setup (small dimensions and low matter flowrates). Since this issue is the core of the article, the emphasis will be placed on the key points allowing this requirement to be met;

including the design of the fast quenching system ensuring that the reaction only occurs inside the reactor. Moreover, pyrolysis of octanoic acid was selected as illustration. Indeed, from the modeling point of view, pyrolysis can be considered as the most basic thermal cracking reaction, and octanoic acid as the simplest departure surrogate of carboxylic acids which play a key role in various industrial applications. These include the steam reforming cracking of waste cooking oil [9], the ethyl biodiesel combustion [10], or the biofuel and bioproduct production from biomass pyrolysis [11-13] or hydrothermal liquefaction [14]. Validity of the kinetic data generated with the designed tubular reactor, including the fast quenching system, will be checked through measurement repeatability and material balances.

2. Experimental Setup Description

A schematic overview of the designed continuous bench-scale experimental setup is given in **Figure 1**. Five major units operating at constant near-atmospheric pressure can be clearly observed: (i) Liquid reagent (octanoic acid) and carrier gas (nitrogen) feed unit; (ii) Mixing and preheating unit of the reactor feed (mixture of reagent and carrier gas); (iii) Gas phase chemical reaction unit; (iv) Multi-stage unit of reaction product trapping: fast quenching, condensation and gas-liquid separation; (v) Reaction product analysis unit. Air supply that was introduced into the feed unit is used at the end of each experiment to quantify the coke formed during the thermal cracking run. Also located at the feed unit level, a water supply was provided for possible extension of the experimental setup to steam cracking or steam reforming applications.

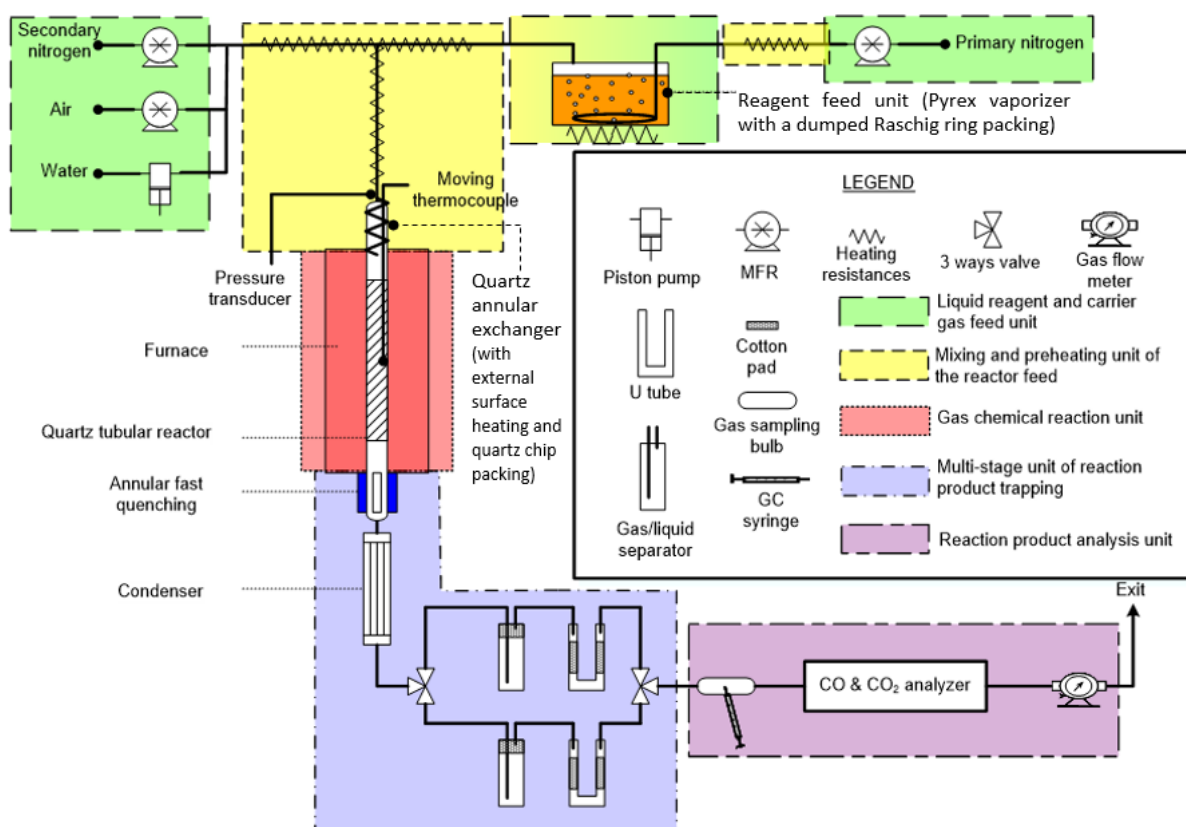


Figure 1. Schematic overview of the designed bench-scale experimental setup.

2.1. Liquid Reagent and Carrier Gas Feed Unit

Two carrier gas feeds were incorporated. The first one named "*primary N₂*" allows for vaporizing partially the liquid reagent to get a homogenous gas phase feed with a constant flowrate. This condition is required to ensure good reproducibility of experiments and to reach steady state during a run (with time independent temperature and concentration profiles throughout reactor). The second carrier gas feed, named "*secondary N₂*", allows for conducting experiments by controlling independently the reagent dilution and the feed residence time. Both carrier gas feeds (primary N₂ and secondary N₂) are supplied via mass flow controllers (maximum volume flowrate at NTP (0°C, 1 atm): 4 NL/min) and are preheated in stainless-steel transfer lines surrounded by heating resistance rolls. The preheated primary N₂ is flowed

at the bottom of the reagent feed unit which is filled with the liquid reagent. The reagent feed unit is a counter-current gas-liquid vaporizer made of pyrex and randomly filled with Raschig rings to enhance mass and heat transfer. A thermally controlled heater system allows for heating it to a constant temperature value (170°C for octanoic acid reagent). The pressure inside the vaporizer is assumed to be equal to the pressure at the inlet of the reactor, the latter being experimentally determined (see [section 2.3.](#)). The exact molar flowrate of the reagent leaving the vaporizer was determined as described in section B1 (**Appendix B, S.I.**) by operating without thermal reaction and collecting, at equally space time intervals, the reagent at the outlet of the reactor.

2.2. Mixing and Preheating Unit of the Reactor Feed

The homogeneous gas phase stream leaving the vaporizer (mixture of primary N₂ and reagent) is mixed with the secondary N₂ in a quartz exchanger of annular geometry (to ensure efficient heat transfer) and filled with quartz chips (to enhance gas-gas mixing and to avoid condensation of reagent). The quartz exchanger (which is integral with the reactor) is surrounded by a roll of heating resistances in order to preheat the reactor feed (mixture of secondary N₂ and of primary N₂ charged with the reagent) to a temperature lower than the reagent thermal cracking temperature (300°C for octanoic acid).

2.3. Gas Phase Chemical Reaction Unit

The unit comprises a quartz tubular reactor which is housed in a thermostatically controlled electric resistance furnace (F 79300-type Thermolyne). For a given furnace temperature set, the reactor temperature profile is measured with a 1000 mm K-type

thermocouple that can be moved along the central axis of the reactor (and of the mixing and preheating reactor feed unit which is integral with the reactor). A pressure transducer, located at the inlet of the reactor, gives a measure of the reactor pressure.

2.4. Multi-Stage Unit of Reaction Product Trapping: Fast Quenching, Condensation and Gas-Liquid Separation

At the outlet of the reactor, the gaseous mixture flow is passed through a quartz annular exchanger the external surface of which is cooled down to a constant temperature (20°C) by mean of a cryostat (F30C-type Compact Julabo with Thermal H5S as coolant). This fast quenching allows for stopping straightaway the reaction as soon as the reacting mixture leaves the reactor. However, due to cooling, the effluents leaving the fast quenching often occur as aerosols (suspension of very fine liquid droplets in a gas flow). Hence, complete separation of the condensable (liquid) from the non-condensable (gaseous) products is a very difficult task which nevertheless needs to be conducted to facilitate the next product analysis step. Thus, in order to achieve the best possible separation of the gas and liquid products, the stream leaving the fast quenching passes through a water-cooled condenser (temperature maintained at 20°C by using the same cryostat as previously). The obtained gas-liquid mixture is sent either to the transient state or to the steady state transfer line which are both identical and linked together at their extremity by two 3-way valves. Each transfer line comprises a glass gas-liquid separator for collecting the liquid product. For trapping the fine liquid droplets that might remain in the gas product, the separator at the outlet is packed with cotton wool. A glass U-tube partially packed with cotton wool was added after the gas-liquid separator as a precaution (but weighing of the cotton wool, before and after each experiment, showed that no product was retained). The gases exiting the gas-liquid separator and U-tube (non-condensable cracking products

mixed in the inert carrier gas) pass through the gas sampler for analysis, and then to the on-line gas flow meter.

2.5. Reaction Product Analysis Unit, Procedure & Material Balances

The pyrolysis products collected as three phases, i.e. gas, liquid, and solid (coke), are analyzed by different techniques as described in **Appendix A (S.I.) [13]**. Molar fractions are typically determined with an accuracy of $\pm 5\%$ for major species and $\pm 10\%$ for minor species. These species are further detailed in **Appendix A** within the context of the present case study focused on octanoic acid pyrolysis. The methodology would however remain the same for another reactant, liquid at ambient conditions.

The operating procedure of the designed bench-scale setup as well as the material balances validating any experiment carried out are described in **Appendix B (S.I.)**.

3. Design of the Tubular Reactor & Fast Quenching System

Sizing of the major units designed in the new experimental setup are listed in **Table 1**. This section focusses on the design of the two units playing a key role in the kinetic data collection: the reactor and the fast quenching system (for which photos are shown in **Figure 2**).

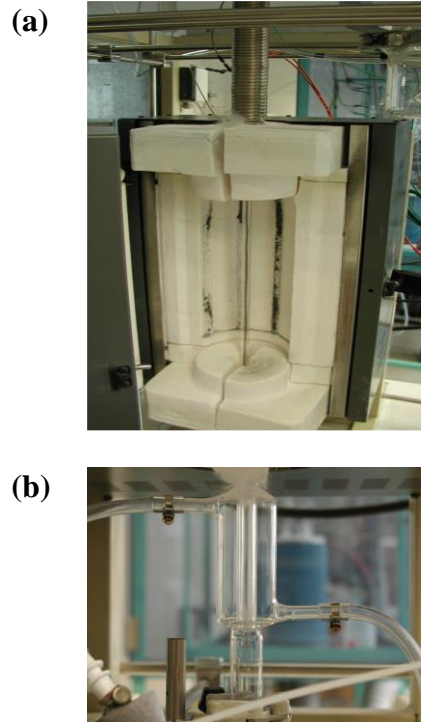


Figure 2. Photos of the designed bench-scale reaction units: **(a)** quartz tubular reactor and its furnace with upstream, the mixing/preheating unit (quartz annular exchanger surrounded by a roll of heating resistances); **(b)** quartz annular fast quenching system (downstream the tubular reactor).

3.1. Reactor Design

In order to ensure flowing conditions close to plug, the tubular reactor should satisfy two criteria (eqs. 1-2):

$$\text{Re} = \frac{u d \rho}{\mu} \geq 10^4 \quad (\text{to guarantee turbulent flow}) \quad (1)$$

$$\frac{L}{d} \geq 10^2 \quad (2)$$

where Re is the Reynolds number whereas u , ρ , and μ stand respectively for the mean real velocity, density, and viscosity of the fluid inside the tubular reactor, for which length and internal diameter are represented by variables L and d. Hence, within the circumstances

encountered in this work (laminar regime imposed by the operating conditions) it should be unlikely that a tubular reactor can be designed with all the requirements to meet rigorous plug flow. As a result, the real objective here was to determine the dimensions of the tubular reactor that would allow approaching as closely as possible plug flow and obtaining then deviations that would be small enough to induce no detectable errors in the kinetic data collection. It is worth mentioning that this objective has good chance to be achieved successfully, although the regime of operating conditions is strongly laminar, because reactions investigated in this work occur in gas phase medium.

Small deviations from plug flow can be considered by using either the dispersion model based on Péclet number (denoted Pe) or the tanks-in-series model based on the equivalent number of CSTRs in series (denoted N). The two models are roughly equivalent and it is thus suggested to use the one which is the most convenient for the user, depending of his purposes [15]. Indeed, for $Pe \geq 50$ or $N \geq 25$, the residence time distributions (RTDs) calculated by the two models for fluid flowing in the reactor lead to symmetric Gaussian curves that are almost superimposable. Also, equating the RTD variances calculated from each model leads to the well-known equivalence relation:

$$Pe = 2(N - 1) \tag{3}$$

This relation is all the more in agreement with experimental observations that Pe and N reach high values (i.e. $Pe > 100$ and $N > 50$). Nevertheless, satisfactory agreement with experiments is obtained with $Pe \geq 50$ (i.e. $N \geq 25$, eq. 3); thus, this value of the Pe criterion was selected to consider a gas-phase laminar tubular reactor very close to plug flow.

In practice, the tanks-in-series model is easier to implement than the dispersion model. In addition, experiments show that the tanks-in-series model can be applied with confidence for non-ideal flows ranging from perfectly mixed flow to plug flow whereas the dispersion model should be devoted to flow close to plug. Also, the tanks-in-series model can be used with any

kinetics [15]. Nevertheless, the dispersion model has the advantage that its key parameter Pe can be easily estimated by correlations involving thermophysical properties of the reacting fluid and resulting from extensive work published in the literature. On the other hand, no estimation method was proposed for the key parameter N . Consequently, the dispersion model was selected for the reactor design of the experimental setup. By contrast, the tanks-in-series model could be used for the modeling and simulation with CHEMKIN II [2] of the kinetic data collected from the designed reactor (the number of equivalent CSTRs in series N would previously be determined for each experiment from the Pe number and eq. 3, and each of the N CSTRs would operate at different temperatures to meet the temperature profile of the tubular reactor induced by the furnace used; see section 4).

Table 1. Dimensions of the experimental setup major units ^a

Units	Features	Variable	Value / m	
Carrier gas and reagent feed	Pyrex vaporizer with			
	a dumped Raschig ring packing	d_i h_{packing}	0.14 0.1	
Mixing and preheating of the reactor feed	Quartz annular exchanger with external surface	L D_i	0.2 0.03	
	heating and quartz chip packing	D_e h_{packing}	0.01 0.03	
	Reactor	Quartz tube	L	0.55
			d_i	0.008
Fast quenching	Quartz annular exchanger with external surface	L D_i	0.1 0.008	
	cooling	d_e	0.006	

^a L : length; d_i : internal diameter; D_i : internal diameter of the annular exchanger large tube; d_e : external diameter of the annular exchanger small tube (D_i and d_e are sizes of the annular exchanger surfaces in contact with the gas flow; **Figure 1**).

3.1.1. Equations Related to the Reactor Sizing

The purpose here is to establish the equations estimating the tubular reactor dimensions (internal diameter d and length L) from the thermophysical properties of the gas reacting fluid. By combining the Péclet number definition (eq. 4) with the correlation proposed in the literature [15] for the axial dispersion coefficient D_{ax} under laminar flow in an empty tube (eq. 5), it was possible to express Pe in terms of d , L , and the properties of the gas reacting fluid that characterize the process: mean real velocity u and molecular diffusion coefficient D_{dif} (eq. 6).

$$Pe = \frac{u L}{D_{ax}} \quad (4)$$

$$D_{ax} = D_{dif} + \frac{u^2 d^2}{192 D_{dif}} \quad (5a)$$

$$\text{applicable when } Re < 2300 \quad (5b)$$

$$\text{and when } L \gg 3d \text{ with } \frac{u \cdot d}{D_{dif}} < \frac{10^2 L}{3 \cdot d}, \quad \text{i.e. } \frac{L}{d} > 3 \cdot 10^{-2} \frac{u \cdot d}{D_{dif}} \quad (5c)$$

$$Pe = \frac{u L}{D_{dif} + \frac{u^2 d^2}{192 D_{dif}}} \quad (6)$$

Then, in order to get an expression as a function of length L , d^2 was deduced from the residence time τ of the stream entering the reactor with a molar composition y_A in reagent A and a total volume flowrate $Q_T^{\text{in}}(T_R, P_R, y_A)$ evaluated at the operating reactor temperature and pressure T_R and P_R (eq. 7).

$$d^2 = \frac{4 Q_T^{\text{in}} \tau}{\pi L} \quad (7)$$

Combining the mean real velocity of the fluid flowing the reactor ($u = L/\tau$) with eqs. 6-7 led to a new expression of Pe involving parameters that can be determined either by estimation methods or by direct measurement on the experimental setup:

$$Pe = \frac{L^2}{D_{\text{dif}} \tau + \frac{Q_T^{\text{in}} L}{48 \pi D_{\text{dif}}}} \quad (8)$$

This expression (eq. 8) is equivalent to a second order equation in L depending only on Pe, D_{dif} , $Q_T^{\text{in}}(T_R, P_R, y_A)$, and τ :

$$L^2 - \frac{Pe Q_T^{\text{in}}}{48 \pi D_{\text{dif}}} L - D_{\text{dif}} \tau Pe = 0 \quad (9)$$

Solving eq. 9 led to the unique acceptable solution with physical meaning:

$$L = \frac{\frac{Q_T^{\text{in}} Pe}{48 \pi D_{\text{dif}}} + \sqrt{\left(\frac{Q_T^{\text{in}} Pe}{48 \pi D_{\text{dif}}}\right)^2 + 4 D_{\text{dif}} \tau Pe}}{2} \quad (10)$$

Finally, given the reactor length L from eq. 10, the reactor internal diameter d could then be determined by eq. 7.

3.1.2. Calculation Procedure and Numerical Solving Related to the Reactor Sizing

The sizing of the tubular reactor requires a few calculation steps as given in the following.

- (1) Choice of an initial value for Pe. The choice $Pe = 50$ was selected since that condition induces a real reactor flow sufficiently close to plug flow (as mentioned previously).
- (2) Selection of operating conditions typically used in pyrolysis process (key variables: τ , T_R , P_R , and $Q_T^{\text{in}}(T_R, P_R, y_A)$) and first evaluation of the reactor length L. These conditions, together with the estimation of the property required for parameter-L calculation (eq. 10), i.e. the molecular diffusion coefficient D_{dif} for the binary system

- [octanoic acid (reagent) – nitrogen (carrier gas)], are listed in **Table 2**. Details of the estimation method used for parameter- D_{dif} are given in **Appendix C [16-19]**. Solving **eq. 9** led to the physical meaning solution $L = 0.49$ m that was rounded to $L = 0.55$ m because of the dimensions of the available furnaces.
- (3) Given the L rounded value obtained in step 2, calculation of the reactor internal diameter d (**eq. 7**). Here again, the value obtained for parameter- d ($d = 8.3 \cdot 10^{-3}$ m) was rounded to the nearest inferior value corresponding to a standard diameter ($d = 8 \cdot 10^{-3}$ m).
 - (4) Checking that the Pe criterion ($Pe \geq 50$) is still satisfied in spite of the rounded values adopted for L and d parameters. Calculation of the mean real velocity for the gas reacting fluid and of the axial dispersion coefficient (u and D_{ax} respectively, **Table 2**) led to $Pe = 56$, allowing to conclude that the flow inside the actual reactor would be close to plug flow under these process operating conditions.
 - (5) Checking the applicability range of the D_{ax} correlation used (i.e. **eq. 5b-c**). Values calculated for Re and the terms involved in **eq. 5c (Table 2)** helped to verify that this last step of L and d calculation procedure was achieved successfully.

3.1.3. Checking the Validity of the Obtained Reactor Sizing for the Whole Experiments Planed in the Work

For the whole pyrolysis experiments planed, the tubular reactor will have to operate typically under $P_R = 1.067$ bar and T_R ranging from 600 to 900°C, with a gaseous feed stream of molar composition 0.05 in reactant (y_A) and of total volume flowrate $Q_T^{in}(T_R, P_R, y_A)$ ranging from 1.5 to 6 L·min⁻¹, the resulting τ ranging from 0.27 to 1 s. These operating conditions lead to values of the Pe criterion included in the interval [46–164]) (with Re in the range [92–205]). One run corresponding to $P_R = 1.067$ bar, $\tau = 0.36$ s, with $T_R = 600^\circ\text{C}$ and

$Q_T^{\text{in}}(T_R, P_R, y_A) = 4.5 \text{ L} \cdot \text{min}^{-1}$ which are extrema (respectively minimum and maximum) operating conditions of the bench-scale setup, led to a slightly low Pe value ($Pe = 46$) but still very close to the limit condition ($Pe \geq 50$). Hence, it is allowed to consider that the designed tubular reactor will be close to a plug flow reactor for the whole experiments planned in the present work.

3.2. Fast Quenching System Design

The objective of the fast quenching system is to cool the reactive gas mixture leaving the reactor almost instantaneously from 600 to 200°C. This assumes that the time required for this operation be much smaller than the residence time of the feed stream inside the reactor. Thus, one could reasonably assume that the thermal reaction process only occurs inside the reactor, and therefore is definitely stopped at the outlet. Within this context, the selected system was a quartz heat exchanger with an annular cross section where the heat would be transferred from the reacting fluid flowing inside the annular space to the outside wall. In practice, the heat exchanger (**Figure 3**) consisted of a central tube enclosed in a double-shell (two concentric pipes) in which the cool-fluid flows counter-currently to the warm-fluid and is maintained at low temperature (20°C; [section 2.4.](#)). Within this geometric configuration, the annular space is then formed by the empty volume between the wall of the central tube (diameter d_e) and the inside wall of the double-shell inner pipe (internal diameter D_i , length L).

Table 2. Pyrolysis operating conditions selected for the reactor sizing (internal diameter d and length L) and the estimation of the thermophysical properties for the gas reacting fluid ^a

Selected Pyrolysis Operating Conditions	
T_R / K	1000
P_R / bar	1.067
y_A	0.05
$Q_E(T_R, P_R, y_A) / (\text{m}^3 \cdot \text{s}^{-1})$	$8.33 \cdot 10^{-5}$
τ / s	0.36
Estimation of the Thermophysical Properties Related to the Reacting Fluid	
$D_{\text{dif}} (1000 \text{ K}, 1.067 \text{ bar}) / (\text{m}^2 \cdot \text{s}^{-1})^b$	5.644
$\rho / (\text{kg} \cdot \text{m}^{-3})^c$	0.488
$\mu / (\text{Pa} \cdot \text{s})^c$	$3.604 \cdot 10^{-5}$
Regime Variables	
$u / (\text{m} \cdot \text{s}^{-1})$	1.657
$D_{\text{ax}} / (\text{m}^2 \cdot \text{s}^{-1})$	$1.627 \cdot 10^{-2}$
Pe	56
Re	179
Reactor Sizing	
L/m	0.49 rounded to 0.55
d/m	$8.3 \cdot 10^{-3}$ rounded to $8 \cdot 10^{-3}$
Checking the Applicability of the D_{ax} Correlation Used	
Re	$179 < 2300$ (laminar flow)
L/d	68.75
$3 \cdot 10^{-2} \frac{u \cdot d}{D_{\text{dif}}}$	7.05 (eq. 5c satisfied)

^a Because products formed during pyrolysis of octanoic acid cannot be known prior experiment, the gas reacting fluid was approximate to the gas mixture feeding the reactor, i.e. octanoic acid (reagent) and nitrogen (carrier gas) with a molar fraction $y_A = 0.05$ in reagent. ^b Details of the estimation method used for evaluating the molecular diffusion coefficient for the binary system [octanoic acid (reagent) – nitrogen (carrier gas)] are given in **Appendix C [16-19]**. ^c Estimation from DIPPR database [16].

With regard to the fast quenching system sizing, when D_i and d_e diameters are given, the issue is to calculate the length L together with the duration t_{op} that is required to achieve the desired cooling stage. These calculations are described in the following section.

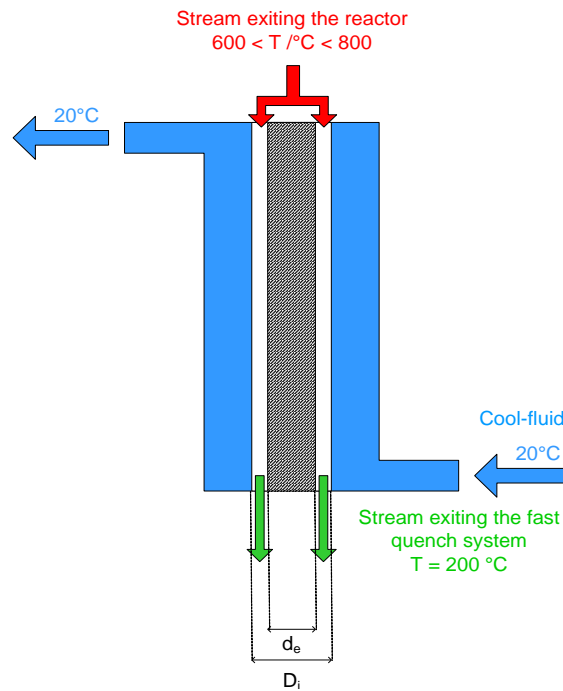


Figure 3. Schematic overview of the fast quenching system (longitudinal section).

3.2.1. Equations Related to the Fast Quenching System Sizing

The equality between the expression of the heat-transfer rate ϕ deriving from the enthalpy balance over the entire exchanger and that deriving from the heating surface area required for the exchanger sizing leads to:

$$\phi = \rho Q \bar{C}_P (T_1 - T_2) = U A \Delta T_m \quad (11)$$

The right-hand side of eq. 11 involves variables characteristic of the gas reacting mixture cooled from temperature T_1 to T_2 (ρ , Q , and \bar{C}_P designate respectively the density, the volume flowrate, and the mass heat capacity at constant pressure of the fluid). All these variables were

determined at the mean temperature $T_m = (T_1 + T_2)/2$ (and at the inlet reactor pressure for ρ and Q). Regarding the left-hand side of eq. 11, variable U designates the overall heat-transfer coefficient based on the inside surface area A of the double-shell inner pipe which is in contact with the warm reacting fluid. The inside surface area A is then given by:

$$A = \pi D_i L, \quad (12)$$

and the overall heat-transfer coefficient U was reasonably approximated to the individual convective heat-transfer coefficient h_G for the warm reacting fluid inside the double-shell inner pipe. Indeed, the individual thermal resistances related to the double-shell inner pipe and to the cool-fluid outside were small enough to be neglected in comparison with the individual thermal resistance of the warm-reacting fluid inside ($1/h_G$). This offered the possibility to equate with sufficient accuracy $U \approx h_G$, allowing h_G estimation from the Nusselt number definition:

$$U \approx h_G = \frac{\lambda Nu}{D_h} \quad \text{with} \quad D_h = (D_i - d_e) = 2\delta \quad (13)$$

In eq. 13, D_h and δ are respectively the hydraulic (or equivalent) diameter and the thickness of the annular space, while λ is the thermal conductivity of the reacting fluid (also calculated at T_m). Combination of Eqs. 11-13 leads to:

$$\rho Q \bar{C}_P (T_1 - T_2) = \frac{\lambda}{2\delta} Nu \cdot \pi D_i L \Delta T_m \quad (14)$$

which can be re-written as following in order to obtain an expression for the fast quenching system length L :

$$L = \frac{2}{\pi} \cdot \frac{Q \rho \delta \bar{C}_P (T_1 - T_2)}{\lambda \Delta T_m D_i Nu} \quad (15)$$

The logarithmic mean temperature difference ΔT_m introduced from eq. 11 is calculated by:

$$\Delta T_m = \frac{T_1 - T_2}{\ln \left(\frac{T_1 - T_{wc}}{T_2 - T_{wc}} \right)} \quad (16)$$

where T_{wc} is the temperature of the double-shell inner pipe wall (cold-fluid side) considered constant and equal to the cooling-fluid temperature (20°C) along the whole fast quenching system length. This assumption agrees with the approximation made for the overall heat-transfer coefficient U . With respect to Nu estimation required in eq. 15, for laminar flow ($Re < 2300$) occurring in annular space, the following correlation is recommended [15]:

$$Nu = \left(1 + 0.14 \sqrt{\frac{D_i}{d_e}} \right) \cdot \left(3.66 + \frac{0.19 Gz^{0.8}}{1 + 0.117 Gz^{0.467}} \right) + 1.2 \left(\frac{D_i}{d_e} \right)^{0.8} - 0.51 \sqrt{\frac{D_i}{d_e}} \quad (17)$$

where the Graetz number Gz is determined from Re (eq. 1 with D_h used instead of d -variable) and the Prandtl number Pr ($Pr = \bar{C}_p \mu / \lambda$) according to:

$$Gz = Re Pr \frac{D_h}{L} \quad (18)$$

For t_{op} estimation, expression of the gas reacting fluid velocity u inside the annular space needs to be firstly determined. Being highly temperature-dependent, the gas reacting fluid velocity u varies considerably from point to point along the fast quenching system length L . Therefore, t_{op} estimation required to start with a differential equation which is then integrated over the entire fast quenching system length. From its physical meaning, t_{op} can be assimilated to the residence time of the reacting fluid inside the fast quenching system that is required to achieve the desired cooling and then can be expressed as:

$$t_{op} = \int_0^L \frac{dx}{u} \quad (19)$$

$$\text{with } u = \frac{4}{\pi} \cdot \frac{Q(T_{NTP}, P_{NTP}, y)}{(D_i^2 - d_e^2)} \cdot \frac{T}{T_{NTP}} \cdot \frac{P_{NTP}}{P} = B \cdot T \quad (20)$$

where B is a constant (isobaric cooling without pressure drop). According to eqs. 15-16, the abscissa x of the fast quenching system where the gas reacting fluid has a temperature T can be expressed as:

$$x = \frac{2}{\pi} \cdot \frac{Q(T_{NTP}, P_{NTP}, y) \rho \delta \bar{C}_P}{\lambda D_i Nu} \ln \left(\frac{T_1 - T_{wc}}{T - T_{wc}} \right) = A \cdot \ln \left(\frac{T_1 - T_{wc}}{T - T_{wc}} \right) \quad (21)$$

where A can be considered as a constant if assuming that Nu is invariant over the entire fast quenching system length. Differentiation of eq. 21 gives:

$$dx = \frac{A}{T - T_{wc}} dT \quad (22)$$

which combined with eqs. 19-20 leads to the desired expression for t_{op} in terms of the operating conditions related to the cooling process and of the thermophysical properties related to the gas reacting fluid exiting the reactor (with a density at normal temperature and pressure conditions ρ_{NTP}):

$$t_{op} = \frac{\bar{C}_P \delta^2 \rho_{NTP} T_{NTP}}{\lambda T_{wc} Nu} \cdot \frac{P}{P_{NTP}} \cdot \frac{D_i + d_e}{D_i} \cdot \ln \left(\frac{T_1 - T_{wc}}{T_2 - T_{wc}} \cdot \frac{T_2}{T_1} \right) \quad (23)$$

Note that assuming Nu invariant in eq. 21, despite its dependency in terms of the fast quenching system x-abscissa, leads to over-estimate t_{op} .

Finally, it is worth mentioning that developments described in this sub-section do not consider radiation heat-transfer. When this phenomenon is occurring, the given equations result in over-estimating the fast quenching system length and the required cooling duration. Nevertheless, most of mono- and di-atomic gases, like nitrogen, have almost no influence on radiation heat-transfer. As a result, developments described above are consistent within the context of the present work.

3.2.2. Calculation Procedure and Numerical Solving Related to the Fast Quenching System Sizing

According to eq. 23, a very thin annular space is recommended to ensure a very fast cooling of the gas reacting fluid (t_{op} proportional to δ^2 , eq. 23). Therefore, an annular space with sizes $D_i = 8$ mm and $d_e = 6$ mm (corresponding to a thickness $\delta = (D_i - d_e) / 2 = 1$ mm) was selected. Furthermore, products formed during pyrolysis of the investigated compound (octanoic acid here) are not known prior reaction. These products are also highly diluted with nitrogen. Therefore, the gas reacting mixture entering the fast quenching system was considered reasonably as pure nitrogen. Parameters required for estimation of the fast quenching system length L and of the cooling duration t_{op} are listed in **Table 3**. Thermophysical properties for nitrogen were determined at $T_m = 673.15$ K and $P_R = 1$ atm. Also, L and t_{op} calculation was made by considering the maximum volume flowrate allowed by the bench-scale experimental setup. Indeed, such an operating condition leads to over-estimate L and t_{op} parameters (eqs. 15 and 23).

The expression obtained for L (eq. 15) involves Nu-number whose expression retained here involves in turn L -parameter via Gz-number (eqs. 17-18). Therefore, an iterative calculation procedure based on the combination of eqs. 15-18 is required to obtain the desired L -parameter value. Results obtained over the successive iterations are listed in **Table 4**. Eq. 23 leads then to the wanted t_{op} -parameter value: $t_{op} = 3.9 \cdot 10^{-3}$ s, i.e. $t_{op} \approx 4$ ms. The obtained cooling time is very short compared to the minimum residence time τ_{min} of the material flow in the reactor ($\tau_{min} \approx 400$ ms). The sizing of the fast quenching performed at this stage is in full agreement with the objectives set and should have no effect on the kinetic data generated. Nevertheless, for technical reasons and to ensure a very efficient cooling even for other applications than the one developed in this work (i.e. for thermal reactions operating at

temperatures higher than 1000°C), a length of 10 cm was finally adopted for the fast quenching system. The corresponding cooling duration of the warm reacting fluid exiting the reactor (with parameters of **Table 3** invariant) is then $t_{op} = 5$ ms, which is still much less than the minimal residence time of the feed stream τ_{min} (≈ 400 ms).

Table 4. Results obtained over the iterative calculation procedure yielding to the wanted length L value of the fast quenching system

Iteration rank k	L_{k-1} / mm	Gz	Nu	L_k / mm
1	5.0	25.145	7.0797	9.1
2	9.1	13.816	6.4627	10.0
3	10.0	12.573	6.3839	10.0

4. Sample of Kinetic Data Generated with the Designed Bench-Scale Setup

The validity of the designed bench-scale PFR, including the fast quenching system, was checked by performing two experiments of octanoic acid pyrolysis under the same operating conditions in order to assess repeatability and material balances. The repeatability was evaluated by considering the standard deviations (SD, eq. 24) and the coefficients of variation (CV, eq. 25) on the operating conditions and on the molar fractions of the reactant (octanoic acid) and of the major pyrolysis products. The material balances were considered in terms of C-H-O atoms (detailed calculations given in **Appendix B, S.I.**).

$$SD = \sqrt{\frac{\sum_{k=1}^{N_p} (v_k - \bar{v})^2}{N_p - 1}} \quad (\text{eq. 24})$$

$$CV(\%) = \frac{SD}{\bar{v}} \cdot 100 \quad (\text{eq. 25})$$

where v_k is the variable (pyrolysis operating conditions or molar fractions of reactant or products) recorded during the k^{th} experiment, the total number of which is N_p leading to an average value \bar{v} .

Regarding the pyrolysis conditions of the experiments, these were selected in order to achieve typical values of three reaction parameters impacting significantly the chemical nature and amount of the products formed: reaction temperature, residence time and extent of dilution of the reactant fluid.

Figure 4 displays the reactor temperature profile measured for the selected oven set point temperature. This temperature profile allows for estimating the average reactor temperature (T_R) used to check the Pe criterion under the actual experiment conditions ($Pe \geq 50$). Moreover, for simulating octanoic acid pyrolysis with the tanks-in-series model, the reactor temperature profile is also required to define the reaction temperature of each of the N CSTRs (with N given by the Pe value calculated at T_R ; [eq. 3, section 3.1](#)).

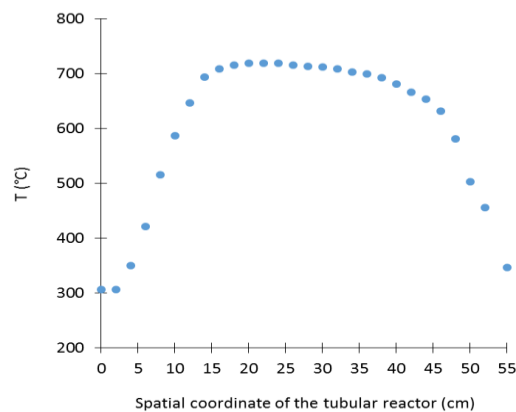


Figure 4. Temperature profile along the tubular reactor for a furnace set point temperature of 700°C.

Results, i.e. the standard deviations and coefficients of variation relating to the operating conditions, the reactant conversion, the material balances and the chemical species molar fractions, are gathered in **Table 5**.

Table 5. Standard deviations (SD) and coefficients of variation (CV) on experimental results generated in this work to validate the designed bench-scale setup – Case study of octanoic acid pyrolysis – Experiments R1 and R2.

Operating conditions	R1	R2	Mean	SD	CV (%)
Furnace set point temperature (°C)	700	700	700	0.0	0.0
Average reactor temperature (°C)	601	601	601	0.0	0.0
Reactor pressure (Torr)	860	862	861	1.4	0.2
Residence time (ms)	422	423	423	0.7	0.2
Pe number	46	46	46	0.0	0.0
Conversion of reactant (%) ^a	7.8	9.2	8.5	0.9	8.5
Material balance in C atom ^a	-0.4	-0.5	-0.5	0.1	0.2
Material balance in H atom ^a	-0.3	-0.3	-0.3	0.0	0.0
Material balance in O atom ^a	1.4	1.7	1.6	0.2	12.5
Chemical species	Molar fractions (solvent-free) %				
Methane (CH ₄)	2.13	2.40	2.27	0.14	6.1
Ethylene (C ₂ H ₄)	7.21	8.23	7.72	0.51	6.6
Propene (C ₃ H ₆)	2.48	2.43	2.46	0.02	1.0
1-Butene (C ₄ H ₈)	0.82	0.87	0.85	0.03	3.0
1-Pentene (C ₅ H ₁₀)	0.54	0.53	0.54	0.01	1.0
1-Hexene (C ₆ H ₁₂)	0.77	0.78	0.78	0.01	0.8
1-Heptene (C ₇ H ₁₄)	0.27	0.25	0.26	0.01	4.9
Acetic acid (CH ₃ COOH)	0.74	0.80	0.77	0.03	3.9
2-Propenoic acid (C ₂ H ₃ COOH)	1.92	2.32	2.12	0.20	9.7
4-Pentenoic acid (C ₄ H ₇ COOH)	0.41	0.44	0.43	0.01	3.5
5-Hexenoic acid (C ₅ H ₉ COOH)	0.54	0.59	0.57	0.02	3.7
6-Heptenoic acid (C ₆ H ₁₁ COOH)	0.23	0.24	0.24	0.01	2.7
Octanoic acid (C ₇ H ₁₅ COOH)	78.42	75.14	76.78	1.64	2.1
Hydrogen (H ₂)	1.42	1.68	1.55	0.13	8.3
Carbon monoxide (CO)	0.53	0.00	0.27	0.26	100.0
Carbon dioxide (CO ₂)	0.84	1.03	0.94	0.09	10.0

^a Molar basis

As it can be observed, the C-H-O element balances are very good. The largest deviations are obtained for the element O because some oxygenated species could not be quantified accurately (such as water, formaldehyde, and ethenone). Regarding the coefficient of variation, very high values are obtained for CO and CO₂. This is very likely due to the analytical method used for quantifying these components (infrared analyzer; **Appendix A [13], S.I.**). Indeed, the

observed levels of CO and CO₂ are in the low detection limits of the equipment (involving thus larger errors of measurements). For the other species, values of the coefficient of variation are inferior to 10 %, which is acceptable for GC analyses. Consequently, it can be said that the experiments for octanoic acid pyrolysis carried out in this work show a rather good repeatability. Such results could not be obtained if the reaction was not occurring under homogenous gas phase, with no wall-effects, through a plug flow reactor with appropriate Pe values, and with an efficient fast quenching system to stop the pyrolysis reaction. Further experiments were generated with the bench-scale setup described in this work to establish a detailed kinetic model for octanoic acid pyrolysis (mechanism and reaction rates) [13] of which predictive performance is illustrated by **Figures 5 and 6**. The detailed chemical kinetic mechanism proposed for octanoic acid pyrolysis [13] was mainly generated from EXGAS software [3-7]. Supplementary reactions, altogether with kinetic and thermodynamic properties of key reactions specific to carboxylic acids, were revisited on the basis of the literature or quantum calculations carried out during the work [13]. Considering the temperature profile along the tubular reactor, the Eulerian approach was selected for modelling the reactor as a succession of CSTRs in series and carrying out the pyrolysis simulations by using the PSR module of CHEMKIN II software [15]. The number of CSTRs was set to 55 in accordance with the number of measurements completed to define the temperature profile of the tubular reactor at a given oven set point (**Figure 4**). For all the supplemental pyrolysis experiments dedicated to the octanoic acid pyrolysis modelling [13], the tubular reactor was close to plug flow according to the Pe criterion ($Pe \geq 50$). The whole of these results confirms the validity of the designed bench-scale setup, particularly regarding the tubular reactor and the fast quenching system.

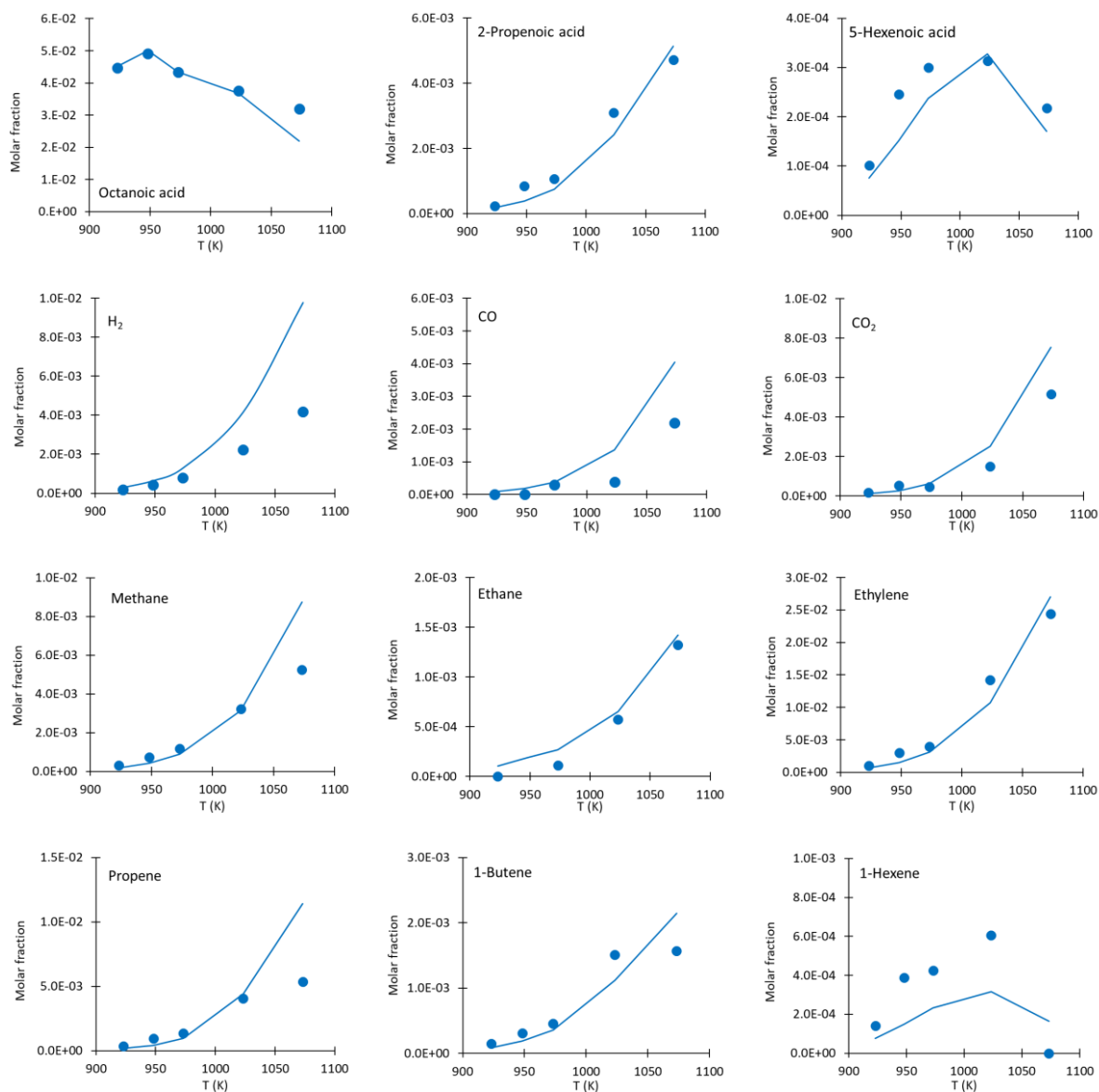


Figure 5. Reactant and key product mole fraction profiles vs. oven set point temperature resulting from octanoic acid pyrolysis carried out in the bench-scale setup designed in this work. The mole fraction profiles obtained for the oven set temperature 973.15 K come from this work (Table 5) while the others come from reference [13].

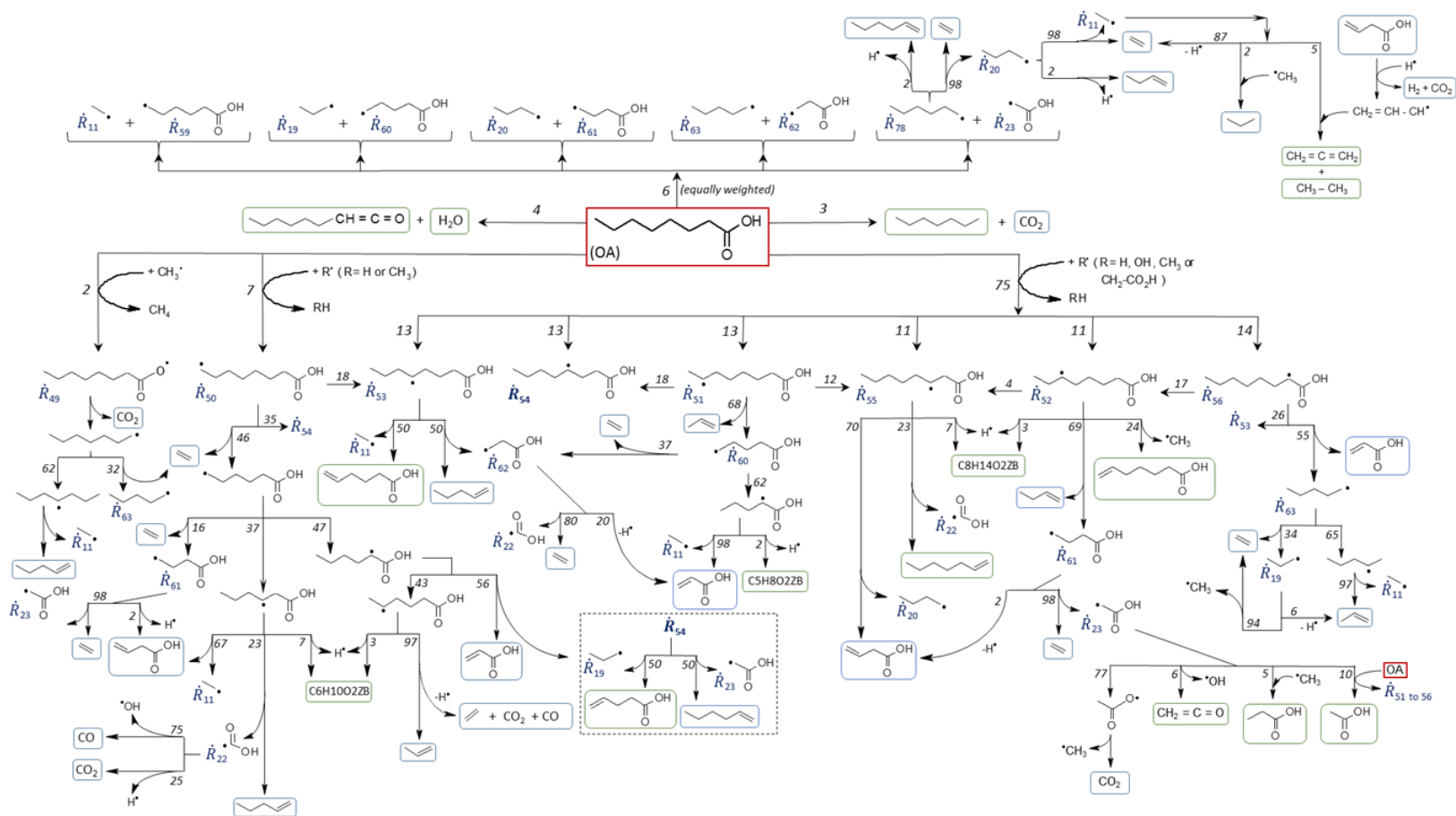


Figure 6. Overall scheme of octanoic acid (OA) pyrolysis as proposed in a previous work [13] on the basis of the experiments generated from the bench-scale setup detailed in this work (operating conditions of pyrolysis: 1 atm, 5% molar OA in N₂, 421 ms residence time, oven set point temperature of 800°C, tubular reactor axial coordinate 22 cm where temperature is 814°C). Italic numbers indicate percent of parent species being converted to daughter species. C₅H₈O₂ZB, C₆H₁₀O₂ZB and C₈H₁₄O₂ZB are globalized species considering isomers from which the major forms are respectively: CH₂=CH-CH₂-CH₂-COOH, CH₂=CH-(CH₂)₃-COOH, and CH₂=CH-(CH₂)₃-COOH.

5. Conclusion

On the basis of a very simple approach, a bench-scale experimental setup devoted to gas phase kinetic data generation in laminar PFR at constant near-atmospheric pressure has been designed. Within the context of octanoic acid pyrolysis, a sample of kinetic data was generated with satisfactory measurement repeatability and material balances, validating the designed bench-scale setup. This result is obtained because the axial dispersion coefficient, in an empty tube, is much smaller in gas-phase than in liquid-phase, allowing a small sized tubular reactor to be close to the plug flow reactor. Such a feature of the axial dispersion coefficient is not well known by the wider public. The material described in this article should be particularly helpful for research projects dealing with pyrolysis, steam cracking or oxidation of liquid biomass (such as waste cooking oil or bio-oil resulting from wood residue thermal-cracking).

References

- [1] Ganley, J.C. A homogeneous chemical reactor analysis and design laboratory: The reaction kinetics of dye and bleach, *Education for Chemical Engineers* 12 (2015) 20-26.
- [2] Kee, R. J., Rupley, F. M., Miller, J. A. Chemkin II. A Fortran Chemical Kinetics Package for the Analysis of a Gas-Phase Chemical Kinetics. Sandia Laboratories Report, SAND 89-8009B, 1993.

- [3] Warth, V., Stef, N., Glaude, P.A., Battin-Leclerc, F., Scacchi, G., Côme, G.M. Computer-Aided Derivation of Gas-Phase Oxidation Mechanisms: Application to the Modeling of the Oxidation of n-Butane, *Combustion and Flame* 114 (1998) 81.
- [4] Glaude, P. A., Battin-Leclerc, F., Judenherc, B., Warth, V., Fournet, R., Côme, G. M., Scacchi, G., Dagaut, P., Cathonnet, M. Experimental and Modelling Study of the Gas-Phase Oxidation of Methyl and Ethyl-tert-Butyl Ethers, *Combustion and Flame* 121 (2000) 345-355.
- [5] Buda, F., Bounaceur, R., Warth, V., Glaude, P. A., Fournet, R., Battin-Leclerc, F. Progress Toward a Unified Detailed Kinetic Model for the Autoignition of Alkanes from C₄ to C₁₀ Between 600 and 1200 K, *Combustion and Flame* 142 (2005) 170-186.
- [6] Vanhove, G., Minetti, R., Touchard, S., Fournet, R., Glaude, P.A., Battin-Leclerc, F. Experimental and Modelling Study of the Autoignition of 1-Hexene / Iso-octane Mixtures at Low Temperature, *Combustion and Flame* 145 (2006) 272-281.
- [7] Moss, J. T., Berkowitz, M., Oehlschlaeger, M. A., Biet, J., Warth, V., Glaude, P. A., Battin-Leclerc, F. An Experimental and Kinetic Modelling Study of the Oxidation of the Four Isomers of Butanol, *Journal of Physical Chemistry A* 112 (2008) 10843-10855.
- [8] Gornay, J., Coniglio, L., Billaud, F., Wild, G. Octanoic Acid Pyrolysis in a Stainless-Steel Tube: What is the Role of the Coke Formed on the Wall? *Journal of Analytical and Applied Pyrolysis* 87 (2010) 78-84.
- [9] Gornay, J., Coniglio, L., Billaud, F., Wild, G. Steam Cracking and Steam Reforming of Waste Cooking Oil in a Tubular Stainless-Steel Reactor with Wall Effects, *Energy and Fuels* 23 (2009) 5663-5676.

- [10] Coniglio, L., Bennadji, H., Glaude, P.A., Herbinet, O., Billaud, F. Combustion Chemical Kinetics of Biodiesel and Related Compounds (Methyl and Ethyl Esters): Experiments and Modeling Advances and Future Refinements, *Progress in Energy and Combustion Science* 39 (2013) 340-382.
- [11] Mohan, D., Pittman, C.U., Steele, P.H. Pyrolysis of Wood/Biomass for Bio-oil: A Critical Review, *Energy and Fuels* 20 (2006) 848-889.
- [12] Ferrando, N., Gedik, I., Lachet, V., Pigeon, L., Lugo, R. Prediction of Phase Equilibrium and Hydration Free Energy of Carboxylic Acids by Monte Carlo Simulations, *The Journal of Physical Chemistry B* 117 (2013) 7123-7132.
- [13] Gornay, J., Glaude, P.A., Billaud, F., Coniglio, L. Experiments and Modeling of Octanoic Acid Pyrolysis in a Plug Flow Reactor, *Journal of Analytical and Applied Pyrolysis* 146 (2020) 104767.
- [14] Gollakota, A.R.K., Kishore, N., Gu, S. A Review on Hydrothermal Liquefaction of Biomass, *Renewable and Sustainable Energy Reviews* 81 (2018) 138-1392.
- [15] Levenspiel, O. *Chemical Reaction Engineering*, 3rd ed., John Wiley & Sons, New York, 1999. ISBN 0-471-25424-X.
- [16] Rowley, R.L., Wilding, W.V., Oscarson, J.L., Knotts, T.A., Iles N.F. DIPPR 801 Data Compilation of Pure Chemical Properties; Design Institute for Physical Properties, AIChE: New York, NY, 2015.
- [17] Wilke, C. R., Lee, C. Y. Estimation of diffusion coefficients for gases and vapors. *Industrial and Engineering Chemistry* 47 (1955) 1253-1257.
- [18] Le Bas, G. *The Molecular Volumes of Liquid Chemical Compounds*, Longmans, Green, New York, 1915.
- [19] Bruce, E. P., Prausnitz, J. M., O'Connell, J. P. *The Properties of Gases and Liquids*, 5th ed., McGraw-Hill, New-York, 2001. ISBN 0-07-011682-2.

SUPPLEMENTARY INFORMATION

Appendix A – Reaction Product Analysis Description

This section gives further details of the products from octanoic acid pyrolysis collected as three phases, i.e. gas, liquid, and solid (coke), together with the techniques selected for their analysis. Operating conditions related to gas chromatography were gathered in **Table A1 [13]**.

A1. Gaseous Products

These are hydrogen (H_2), carbon monoxide (CO), carbon dioxide (CO_2), and hydrocarbons having a carbon number inferior to 4 (C_1 - C_4 cut). All of them were analyzed continuously during the experiment by on-line (for CO and CO_2) or off-line (for H_2 and C_1 - C_4 cut) equipment. Also, standard gas mixtures were used for calibration (and identification concerning the C_1 - C_4 cut). CO and CO_2 were analyzed by a non-dispersive infrared (NDIR) analyzer (Cosma Cristal 300), while H_2 and the C_1 - C_4 cut were analyzed by two gas chromatographs (GC) equipped with different detectors for suitable analysis: a thermal conductivity detector (TCD) for H_2 and a flame ionization detector (FID) for the C_1 - C_4 cut. This equipment allows quantification of CO, CO_2 , and H_2 in terms of percent volume fractions in the whole gas product (equivalent to percent molar fractions with regards to the ideal gas law), while quantification of the C_1 - C_4 cut components is obtained in terms of percent mass fractions in the gas product excluding CO, CO_2 , H_2 , and the inert carrier gas.

A2. Liquid Products

Usually, the thermal cracking liquid product contains a wide variety of components with various functional groups. Therefore, a gas chromatography-mass spectrometry (GC-MS) technique was primarily used for component identification (which includes matching of retention times against known standard compounds under identical GC conditions together with analysis of GC mass spectra thanks to spectral libraries). Nevertheless, for some experiments, a small fraction of components could not be identified using the available GC-MS database and were termed "unidentified" fraction. Quantification of the identified liquid components (i.e. the condensable cracking products and the non-converted octanoic acid) was achieved by using two GC-FID associated with the widely used internal standard method (internal standards selected: n-octane for the condensable products and n-dodecane for octanoic acid).

A3. Coke

Thermal cracking involves coke formation. Coke deposited by successive layers on the reactor walls has to be quantified to account for it in the mass balance. Also, decoking the internal surface of the reactor is necessary to have a good reproducibility of the experiments (constant cross-sectional area of the tubular reactor from one experiment to another) as well as to avoid plugging (and other problems in downstream unit operations). The amount of coke deposit was determined by complete combustion (eq. A1) and analyzing the resulting CO and CO₂ emissions. Details of the procedure are given in **Appendix B** (section B1).



Table A1. Conditions of gas chromatography analyses [13]

Conditions	H ₂	Non-condensable gas (excluding H ₂ , CO, and CO ₂)	Condensate ^a (excluding octanoic acid)	Condensate ^a (exclusively octanoic acid)
GC-type	Intersmat IGC 11-type	Schimadzu GC 17A	Schimadzu GC 17A	Stang ST 200
Detector	TCD	FID	FID	FID
Detector temperature /°C	-	310	310	280
Injector temperature /°C	-	300	300	270
Carrier gas (N ₂) volume flowrate (NmL/min) or linear velocity (cm/s)	30 NmL/min	15 cm/s	15 cm/s	0.91 NmL/min
Split ratio	-	1/100	1/100	1/160
Column	Silica-gel filled column / molecular sieve (5Å thickness) 5 m length, 6 mm inner diameter	Bonded non-polar (methyl silicone) capillary column (P.O.N.A.-type, Hewlet Packard) 50 m, 0.21 mm, 0.5 µm film thickness	Bonded non-polar (methyl silicone) capillary column (P.O.N.A.-type, Hewlet Packard) 50 m, 0.21 mm, 0.5 µm film thickness	Polyethyleneglycol capillary column (CP-Wax 52CB, Chrompack) 50 m, 0.32 mm, 0.2 µm film thickness
Oven temperature program	60°C	60°C (4 min), 60-180°C (10°C/min, 10 min), 180-300°C (10°C/min)	60°C (20 min), 60-300°C (2°C/min, 40 min)	60°C (10 min), 60-240°C (10°C/min, 32 min)
Column pressure program	-	96 kPa (4 min), 96-153 kPa (2.4 kPa/min, 10 min), 153-182 kPa (2.4 kPa/min)	124 kPa (20 min), 124-182 kPa (0.5 kPa/min, 44 min)	-

^a The GC-FID system was calibrated by determining a carbon response factor compared with the internal standard (IS) both for each identified product and for the unconverted reagent (octanoic acid). To cover a wider range of retention times, two IS were used. As previous tests showed that neither n-octane nor n-dodecane were formed during octanoic acid pyrolysis (under the investigated reaction conditions), n-octane was selected as IS for all observed products whereas n-dodecane was used as IS to quantify the non-converted octanoic acid fraction.

Appendix B – Designed Bench-Scale Setup Procedure & Material Balances

This section is dedicated to the operating procedure of the designed bench-scale setup together with the material balances selected as criteria for validating any experiment generating pyrolysis kinetic data.

B1. Operating procedure

Before each run, the exact partial molar flowrate of reagent (octanoic acid) at the inlet of the reactor (F_A^{in}) was determined by operating without thermal reaction. Hence, while the vaporizer was maintained at 170°C (temperature determined during previous tests as leading to vaporization of a sufficient amount of octanoic acid), the reactor together with the mixing and preheating unit were heated firstly at 180°C to avoid thermal cracking as well as condensation of octanoic acid. Then, the liquid product collected in the gas-liquid separator could be assumed to be exclusively pure octanoic acid (assumption verified by GC measurement) and the mass of liquid product m_A^{in} collected at equally spaced time intervals Δt_A (determined by double weighting of the gas-liquid separator including the padding cotton at $\Delta t_A = 10 \text{ min}$) allowed experimental determination of F_A^{in} according to:

$$F_A^{\text{in}} = \frac{m_A^{\text{in}}}{M_A} \cdot \frac{1}{\Delta t_A}, \quad (\text{B1})$$

where M_A is the molecular weight of octanoic acid. Also, it is worth mentioning that m_A^{in} was considered for F_A^{in} determination once steady state was reached (i.e. when m_A^{in} led to a time independent value for F_A^{in}). Also, the mole fraction of octanoic acid in the reactor inlet stream can be determined by:

$$y_A^{\text{in}} = \frac{F_A^{\text{in}}}{F_A^{\text{in}} + F_{N_2}^{\text{in}}} \quad (\text{B2})$$

$$\text{with } F_{N_2}^{\text{in}} \approx F_{N_2}^{\text{out}} = Q_{N_2}^{\text{in}}(\text{NTP}) \cdot \frac{P_{\text{NTP}}}{RT_{\text{NTP}}} \quad (\text{B3})$$

where $Q_{N_2}^{\text{in}}(\text{NTP})$ is the nitrogen volume flowrate supplied via mass flow controller at normal temperature and pressure ($T = 273.15 \text{ K}$ and $P = \text{atm}$). Due to the very low solubility of N_2 in octanoic acid, the partial molar flowrate of N_2 could be considered invariant all along the experimental setup, from the carrier gas feed unit ($F_{N_2}^{\text{in}}$) to the reaction product analysis unit.

After this preliminary stage aimed at determining experimentally the exact partial molar flowrate of reagent entering the reactor, the vaporizer was maintained at constant temperature (170°C) and the desired operating conditions were set: volume flowrate of secondary N_2 (preheated to 170°C) if required to increase the dilution or residence time of the feed, temperature of the mixing and preheating unit of the reactor feed (300°C for all runs), and temperature of the furnace.

Then, took place the experimental setup transient state, the end of which was determined when all operating conditions (reactor pressure and temperatures of the different units) as well as H_2 , CO , and CO_2 emissions (percent fractions measured every 2 minutes) were stationary (actually, the $CO\&CO_2$ NDIR analyzer as well as the H_2 GC were the two key fittings for monitoring the steady-state operation). When the end of the transient state was reached, the reactant flow exiting the condenser was directed through the steady state transfer line, giving the starting point of the experiment. Each run was carried out for a period of 30 to 60 minutes, during which the liquid product was collected (in the gas-liquid separator for weighing and GC analysis at the end of the run) while the gas product was continuously analyzed (with the on-line NDIR analyzer and the off-line GCs) at equally spaced time intervals (5 min for CO , CO_2 , and H_2 , 15 min for the other non-condensable components of the gas product). Gas product

sampling for off-line GC analysis was carried out using a gas bulb located after the gas-liquid separation unit.

At the end of the experiment, the gas effluent leaving the reactor was directed through the transitory transfer line (thanks to the 3-way valves), and the gas-liquid separator was removed from the setup for further analysis.

Then decoking was carried out as follows. After each reaction run, the reactor was heated at very high temperature (furnace temperature: 900°C) while a mixture of nitrogen and air (0.3 and 0.7 NL/min, respectively) was passed through it to burn off the coke formed during the run. In order to optimize decoking, it was decided to proceed by successive oxidation cycles followed by integration on the entire oxidation duration. Therefore, the air volume flowrate was increased by stages at the end of each oxidation cycle while maintaining constant the total volume flowrate of the N₂ and air mixture. For each decoking cycle, percent volume fractions of CO and CO₂ ($\mathcal{Y}_{\text{CO},k}^{\text{coke}}$ % and $\mathcal{Y}_{\text{CO}_2,k}^{\text{coke}}$ %, respectively) were read on the NDIR analyzer at equally spaced time intervals ($\Delta t_k = 2$ minutes) while the total volume of the gas product (CO, CO₂, N₂, and air) was read on the gas meter at ambient temperature and pressure ($V_k^{\text{coke}}(T_a, P_a)$). When the decreasing $\mathcal{Y}_{\text{CO},k}^{\text{coke}}$ % and $\mathcal{Y}_{\text{CO}_2,k}^{\text{coke}}$ % parameters reached constant values, the air volume flowrate was again increased for a new oxidation cycle at constant total volume flowrate (for example, the same value 0.5 NL/min for air and N₂). When the $\mathcal{Y}_{\text{CO},k}^{\text{coke}}$ % and $\mathcal{Y}_{\text{CO}_2,k}^{\text{coke}}$ % parameters remained constant, close to zero, in spite of increasing the air input, all the coke deposit was assumed to be removed from the inner surface of the reactor and the decoking ended. Integration, over all the decoking duration, of $\mathcal{Y}_{\text{CO},k}^{\text{coke}}$ %, $\mathcal{Y}_{\text{CO}_2,k}^{\text{coke}}$ % and $V_k^{\text{coke}}(T_a, P_a)$ parameters recorded during each oxidation cycle yields the mass of coke deposit m_{coke} according to eq. B4 where the gas product was assumed to follow the ideal gas law.

$$m_{\text{coke}} = \frac{P_a}{RT_a} \cdot \left\{ \sum_k V_k^{\text{coke}}(T_a, P_a) \cdot (y_{\text{CO},k}^{\text{coke}} + y_{\text{CO}_2,k}^{\text{coke}}) M_C \right\} \quad (\text{B4})$$

where $y_{j,k}^{\text{coke}}$ is the mole fraction of specie j (CO or CO₂) measured at time k (equal to the volume fraction $\mathcal{G}_{j,k}^{\text{coke}}$ according to the ideal gas law) and M_C is the carbon molecular weight. Finally, the mass of the liquid product collected during each run was determined by weighing of the gas-liquid separator (including padding cotton) at the beginning and the end of the experiment while composition was obtained by GC analysis.

B2. Material Balances

Each experiment was validated by calculating balances in terms of C-H-O element mole number between the inlet and the outlet of the reactor. The C-H-O element mole balances were calculated given the pressure and temperature of the reactor, the volume flowrates of primary and secondary N₂ (at normal temperature and pressure, i.e. 273.15 K and 1 atm), the duration of each run, together with other variables measured along the total run period: • mass of octanoic acid entered into the reactor (m_A^{in}), • mass of the liquid product collected in the gas-liquid separator (m_L^{out}), • mass of the coke deposit (m_{coke}), • mass of the gas product (obtained by: $m_{\text{incond}}^{\text{out}} = m_A^{\text{in}} - m_L^{\text{out}} - m_{\text{coke}}$ thanks to the observation that no material leak was detected during the run); these measurements were of course completed with the percent mole fractions of the identified components in the gas phase and the mass of the identified organic components in the liquid product representing the further termed organic liquid phase (OLP) product, in addition of the molecular formulae of all species formed. The main general equations used to check the C-H-O element mole balances for each run are given in **Table B1**.

Table B1. General equations related to the C-H-O element mole balances made on the designed experimental setup ^a

Equations Related to the Inlet of the Reactor	
• For the Reagent Liquid Feed (Octanoic Acid)	
$F_A^{\text{in}} = \frac{m_A^{\text{in}}}{M_A} \cdot \frac{1}{\Delta t_A}$	(B1)
• For the Carrier Gas Feed (Nitrogen)	
$F_{N_2}^{\text{in}} = \frac{P_{NTP} \cdot Q_{N_2}^{\text{in}}(T_{NTP}, P_{NTP})}{RT_{NTP}}$	(B3)
• Residence Time of the Feed	
$\tau = V_R / Q_T^{\text{in}}(T_R, P_R, y_A)$	(B5)
with $Q_T^{\text{in}}(T_R, P_R, y_A) = F_T^{\text{in}} \cdot \frac{RT_R}{P_R}, y_A = F_A^{\text{in}} / F_T^{\text{in}}$	(B2)
$F_T^{\text{in}} = F_A^{\text{in}} + F_{N_2}^{\text{in}}, \text{ and } F_{N_2}^{\text{in}} = F_{\text{primary } N_2}^{\text{in}} + F_{\text{secondary } N_2}^{\text{in}}$	(B6)
Equations Related to the Outlet of the Reactor ^b	
• Partial Molar Flowrate for Organic Liquid Phase (OLP) Components	
- Weight and Molar Fractions	
$F_{i,\text{cond}}^{\text{out}} = \frac{m_{i,\text{cond}}^{\text{out}}}{M_{i,\text{cond}}^{\text{out}} \cdot \Delta t_{\text{exp}}}$	(B7)
with $m_{i,\text{cond}}^{\text{out}} = m_E \cdot K_{i/E} \cdot \frac{A_i}{A_E}$	
$\bar{x}_{i,\text{cond}}^{\text{out}} = \frac{m_{i,\text{cond}}^{\text{out}}}{\sum m_{i,\text{cond}}^{\text{out}}}$	(B8)
and $x_{i,\text{cond}}^{\text{out}} = \frac{\bar{x}_{i,\text{cond}}^{\text{out}} / M_{i,\text{cond}}^{\text{out}}}{\sum (\bar{x}_{i,\text{cond}}^{\text{out}} / M_{i,\text{cond}}^{\text{out}})}$	
• Intermediate Normalization of the Detected Incondensable Components (Gaseous Products)	
$y_j / \sum y_j$ where j is either an on-line or an off-line quantified product.	(B9)
For an on-line quantified product: $y_j = y_{j,\text{online } G}^{\text{out}} = \mathcal{S}_{j,\text{online } G}^{\text{out}}$	
For an off-line quantified product: $y_j = y_{j,\text{offline } G}^{\text{out}} = \frac{\bar{y}_j / M_j}{\sum \bar{y}_j / M_j}$ with $\bar{y}_j(\%) = \frac{A_j}{\sum A_j} \cdot 100$	
	(B10)

Table B1. Continued ^aEquations Related to the Outlet of the Reactor ^b

• Total Mass of All Incondensable Products

$$m_{\text{incond}}^{\text{out}} = m_{\text{A}}^{\text{in}} - m_{\text{L}}^{\text{out}} - m_{\text{coke}} \quad (\text{B11})$$

• Partial Molar Flowrate for the On-line Detected Incondensable Products (H₂, CO, and CO₂)

$$\bar{y}_{j,\text{online G}}^{\text{out}} = \frac{y_{j,\text{online G}}^{\text{out}} \cdot M_{j,\text{online G}}^{\text{out}}}{\left(\sum M_{j,\text{online G}}^{\text{out}} \cdot y_{j,\text{online G}}^{\text{out}} \right) + \left(1 - \sum y_{j,\text{online G}}^{\text{out}} \right) \cdot M_{\text{N}_2}} \quad \text{where } y_{j,\text{online G}}^{\text{out}} = \mathfrak{g}_{j,\text{online G}}^{\text{out}} \quad (\text{B12})$$

$$F_{j,\text{online G}}^{\text{out}} = \frac{m_{j,\text{online G}}^{\text{out}}}{M_{j,\text{online G}}^{\text{out}} \cdot \Delta t_{\text{exp}}} \quad \text{with } m_{j,\text{online G}}^{\text{out}} = (m_{\text{incond}}^{\text{out}} + F_{\text{N}_2}^{\text{in}} \cdot M_{\text{N}_2} \cdot \Delta t_{\text{exp}}) \cdot \bar{y}_{j,\text{online G}}^{\text{out}} \quad (\text{B13})$$

• Partial Molar Flowrate for the Off-line Detected Incondensable Products (Other Gases than H₂, CO, and CO₂)

$$m_{j,\text{offline G}}^{\text{out}} = m_{\text{incond}}^{\text{out}} \cdot y_{j,\text{offline G}}^{\text{out}} \quad (\text{B14})$$

$$F_{j,\text{offline G}}^{\text{out}} = \frac{m_{j,\text{offline G}}^{\text{out}}}{M_{j,\text{offline G}}^{\text{out}} \cdot \Delta t_{\text{exp}}} \quad (\text{B15})$$

• Partial Molar Flowrate for Components Encountered both in the Gas Phase and in the Organic Liquid Phase

$$F_j^{\text{out}} = F_{j,\text{offline G}}^{\text{out}} + m_{i,\text{cond}}^{\text{out}} / (M_{i,\text{cond}}^{\text{out}} \cdot \Delta t_{\text{exp}}) \quad (\text{B16})$$

• Final Normalization of all Detected Component Molar Fractions as a Whole

$$y_k(\%) = (F_k^{\text{out}} / \sum F_k^{\text{out}}) \cdot 100 \quad (\text{B17})$$

$$\text{with } F_k^{\text{out}} = F_{j,\text{online G}}^{\text{out}} \quad \text{for on-line detected gaseous products,} \quad (\text{B18})$$

$$F_k^{\text{out}} = F_{j,\text{offline G}}^{\text{out}} \quad \text{for off-line detected gaseous products,} \quad (\text{B19})$$

$$F_k^{\text{out}} = F_{i,\text{cond}}^{\text{out}} \quad \text{for the organic liquid phase products,} \quad (\text{B20})$$

$$F_k^{\text{out}} = \frac{m_{\text{coke}}}{M_{\text{C}} \cdot \Delta t_{\text{exp}}} \quad \text{for the coke deposit (with } M_{\text{C}} = 12.01145), \quad (\text{B21})$$

and F_k^{out} given by eq. B16 for components encountered both in the organic liquid phase and in the gas phase.

Table B1. Continued ^a

Carbon-Hydrogen-Oxygen Element Mole Balance Equations (%)

$$\frac{(F_{\text{element}}^{\text{in}} - F_{\text{element}}^{\text{out}})}{F_{\text{element}}^{\text{in}}} \cdot 100 \quad (\text{B22})$$

with $F_{\text{element}}^{\text{in}} = v_{\text{element}}^{\text{A}} \cdot F_{\text{A}}^{\text{in}}$, $F_{\text{element}}^{\text{out}} = v_{\text{element}}^{\text{k}} \cdot F_{\text{k}}^{\text{out}}$, and element = C, H, or O.

^a The meaning of each variable is described in the nomenclature section below. ^b The partial molar flowrate of N₂ between the inlet and the outlet of the reactor is invariant. The total mass of coke deposited is given by [eq. B4 \(section B1\)](#).

Nomenclature

A_{E}	Chromatographic area of internal standard E
A_{i}	Chromatographic area of component i
F_{A}^{in}	Inlet octanoic acid molar flowrate
$F_{\text{element}}^{\text{in}}$	Inlet molar flowrate of the considered element
$F_{\text{N}_2}^{\text{in}}$	Inlet nitrogen molar flowrate
$F_{\text{primary N}_2}^{\text{in}}$	Inlet molar flowrate of the primary carrier gas feed (for partial vaporization of the liquid reagent)
$F_{\text{secondary N}_2}^{\text{in}}$	Inlet molar flowrate of the secondary carrier gas feed (for individual control of τ and reagent dilution)
F_{T}^{in}	Inlet total molar flowrate
$F_{\text{element}}^{\text{out}}$	Outlet molar flowrate of the considered element
$F_{\text{l,cond}}^{\text{out}}$	Molar flowrate of the outlet condensable component i
$F_{\text{j}}^{\text{out}}$	Molar flowrate of product j encountered both in the gas phase and in the organic liquid phase
$F_{\text{j,offlineG}}^{\text{out}}$	Molar flowrate of the off-line detected incondensable product j in the outlet gas stream
$F_{\text{j,onlineG}}^{\text{out}}$	Molar flowrate of the on-line detected incondensable product j in the outlet gas stream

$K_{i/E}$	Carbon response factor of component i compared with internal standard E for quantification by GC-FID analysis
m_A^{in}	Inlet mass of octanoic acid collected at equally spaced time intervals $\Delta t_A = 10 \text{ min}$
m_{coke}	Mass of coke deposited during each experiment
m_E	Mass of internal standard E
$m_{i,\text{cond}}^{\text{out}}$	Mass of the outlet condensable component i
$m_{\text{incond}}^{\text{out}}$	Total mass of the detected incondensable products taken as a whole in the outlet gas stream
$m_{j,\text{online G}}^{\text{out}}$	Mass of the on-line detected incondensable product j in the outlet gas stream
$m_{j,\text{offline G}}^{\text{out}}$	Mass of the off-line detected incondensable product j in the outlet gas stream
m_L^{out}	Mass of the outlet organic liquid phase product
M_j	Molecular weight of component j
$M_{i,\text{cond}}^{\text{out}}$	Molecular weight of the outlet condensable component i
$M_{j,\text{offline G}}^{\text{out}}$	Molecular weight of the off-line detected incondensable product j present in the outlet gas stream
$M_{j,\text{online G}}^{\text{out}}$	Molecular weight of the on-line detected incondensable product j present in the outlet gas stream
P_{NTP}	Pressure related to normal temperature and pressure conditions ($P_{\text{NTP}} = 1 \text{ atm}$)
P_R	Pressure at the inlet of the reactor (assumed constant until the outlet)
$Q_{\text{N}_2}^{\text{in}}(T_{\text{NTP}}, P_{\text{NTP}})$	Inlet volume flowrate related of nitrogen at T_{NTP} and P_{NTP}
$Q_T^{\text{in}}(T_R, P_R, y_A)$	Total inlet volume flowrate at T_R and P_R for a molar fraction y_A in reactant A (octanoic acid here)
R	Ideal gas constant ($R = 8.314411 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$)

T_{NTP}	Temperature related to normal temperature and pressure conditions ($T_{\text{NTP}} = 273.15 \text{ K}$)
T_{R}	Average reactor temperature
V_{R}	Reactor volume
$x_{i,\text{cond}}^{\text{out}}$	Molar fraction related to the outlet condensable component i
$\bar{x}_{i,\text{cond}}^{\text{out}}$	Mass fraction related to the outlet condensable component i
y_j	Molar fraction of component j in the gas phase
\bar{y}_j	Mass fraction of component j in the gas phase
$y_{j,\text{offlineG}}^{\text{out}}$	Molar fraction of the off-line detected incondensable product j in the outlet gas stream
$\bar{y}_{j,\text{onlineG}}^{\text{out}}$	Mass fraction of the on-line detected incondensable product j present in the outlet gas stream

Greek symbols

Δt_{A}	Time intervals during which reactant A (octanoic acid) is collected in order to determine its molar flowrate ($\Delta t_{\text{A}} = 10 \text{ min}$)
Δt_{exp}	Duration of experiment
v_{element}^k	Stoichiometric coefficient of “element” in the chemical formula of component k
$v_{\text{element}}^{\text{A}}$	Stoichiometric coefficient of “element” in the chemical formula of octanoic acid
$\mathcal{S}_{j,\text{onlineG}}^{\text{out}}$	Volume fraction of the on-line detected incondensable product j in the outlet gas stream
τ	Residence time of the feed

Appendix C - Estimation Method for the Molecular Diffusion Coefficient D_{dif} – Application to Low Pressure Binary Gas-Phase Diffusion

This section covers briefly the estimation method used in this work for representing diffusion in the reacting mixture that flows along the designed tubular reactor. It should be recalled that because pyrolysis products could not be known before experiments, the binary system feeding the reactor, i.e. [octanoic acid (reagent) – nitrogen (carrier gas)] was considered for representing the diffusion in the reacting mixture. Also, the designed tubular reactor is devoted to kinetic collection at constant near-atmospheric pressure.

Various semi-empirical methods were proposed in the literature for estimating the diffusion coefficients for binary gas systems at low pressures. A modified version of the method by Wilke and Lee [17] was adopted in this work. This version generally agrees with experimental values to 5 % (although discrepancies of around 20 % are possible for binary systems involving polar molecules with long carbon chain). The expression proposed for the diffusion coefficient D_{dif} of the binary mixture of A and B is:

$$D_{\text{dif}} = 7.28 \cdot 10^{-8} \frac{m \cdot (4.340 - m) \cdot T^{3/2}}{P \cdot (V_A^{1/3} + V_B^{1/3})^2 \cdot F(z)} \quad (\text{C1})$$

$$\text{where } m = \sqrt{\frac{1}{M_A} + \frac{1}{M_B}}, \quad F(z) = \left(\frac{0.072}{z^{4.12}} + \frac{0.0062}{z^{1.25}} \right)^{0.125}, \quad \text{and } z = \frac{T}{0.77 \sqrt{T_{C_A} + T_{C_B}}} \quad (\text{C2})$$

Variables T and P are respectively temperature (in K) and pressure (in bar), whereas V_A and V_B are molar volume (in $\text{cm}^3 \cdot \text{mol}^{-1}$) of pure liquid components A and B at their normal boiling point. In these equations, pure components A and B are also characterized by their molecular weights M_A and M_B (in $\text{g} \cdot \text{mol}^{-1}$) and their critical temperatures T_{C_A} and T_{C_B} (in K).

For the estimation of octanoic acid pure liquid molar volume at its normal boiling point, the method by Le Bas [18,19] was adopted. Based on the group contribution concept, the method is very simple and yields a mean deviation of 4 % for the whole of the compounds investigated by the authors.

Calculation of the molecular diffusion coefficient is briefly detailed in the following at $T = 1000$ K and $P = 1.067$ bar corresponding to a set of octanoic acid pyrolysis operating temperature and pressure. By denoting octanoic acid A (reagent) and nitrogen B (carrier gas), the required properties are: $M_A = 144.214$ g · mol⁻¹, $M_B = 28.013$ g · mol⁻¹, $T_{C_A} = 694.26$ K, $T_{C_B} = 126.20$ K, and $V_B = 31.2$ cm³ · mol⁻¹ (DIPPR database [16]). For component A (C₈H₁₆O₂), the method by Le Bas led to:

$$V_A = 8(14.8) + 16(3.7) + 2(12.0) = 201.6 \text{ cm}^3 \cdot \text{mol}^{-1}.$$

Thus, using eq. C2, we obtain: $m = 0.20647$, $z = 4.3875$, and $F(z) = 0.42862$, leading finally with eq. C1 to: $D_{\text{dif}} = 5.644 \cdot 10^{-5} \text{ m}^2 \cdot \text{s}^{-1}$.