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## An introduction to atmospheric chemistry

Max McGillen

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# **An introduction to atmospheric chemistry**

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France

# Lecture outline

- Why does chemical composition matter?
  - Air quality
  - Climate
- Getting to know the atmosphere
  - Physical attributes (structure, temperature, pressure)
  - Chemical composition
- Chemical reactions in the troposphere
- Chemical reactions in the stratosphere
- Approaches to studying the atmosphere

# Part 1 – Why does chemical composition matter?

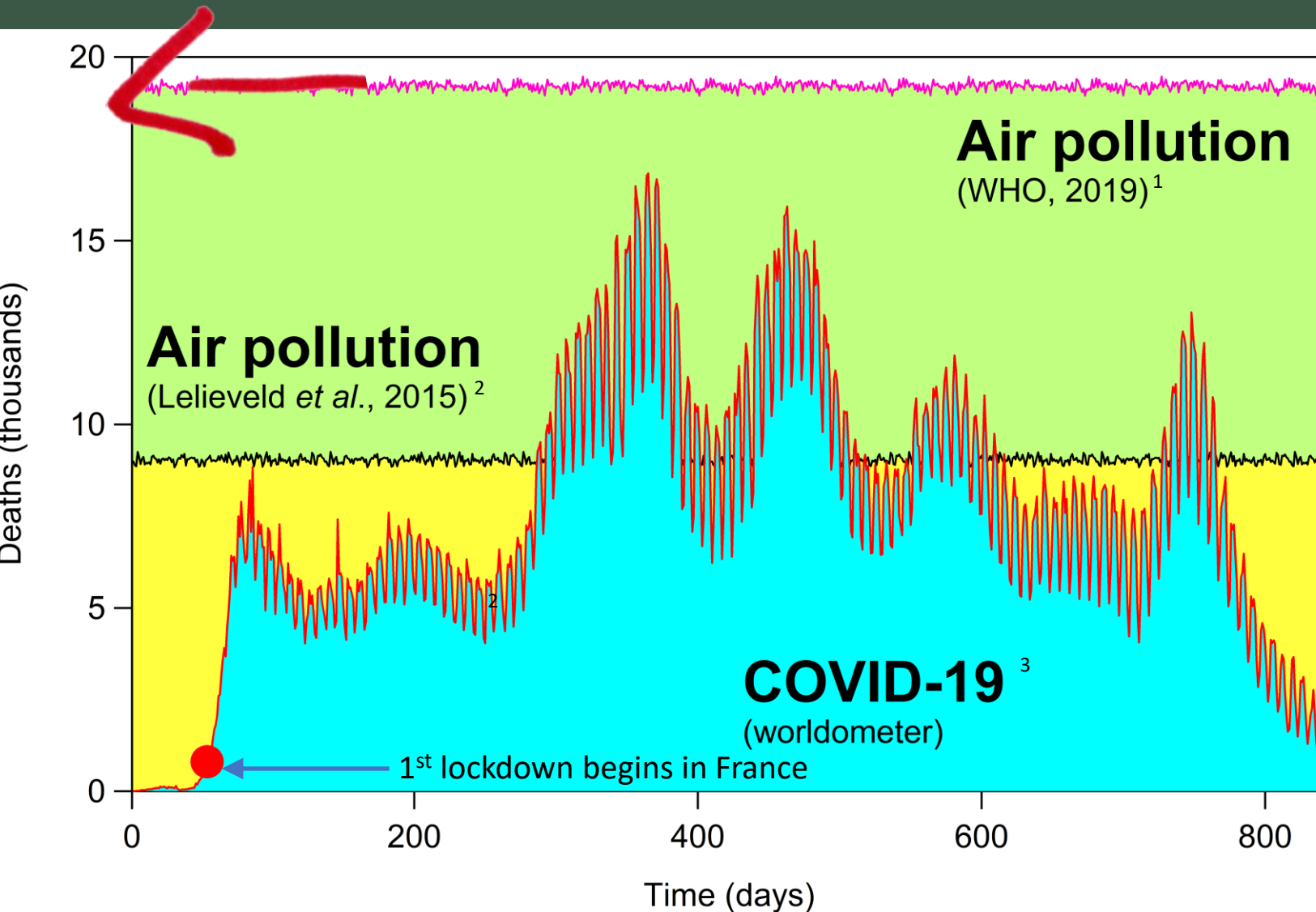
**“the simple act of breathing is killing 7 million people a year and harming billions more”**

Dr. T.A. Ghebreyesus, head of the World Health Organization





# Air pollution: an invisible pandemic?



- Impact of our emissions is defined by atmospheric chemistry
- Air pollution is atmospheric chemistry
- Air pollution is an urgent problem
- And it will get worse in the future<sup>2</sup>
- The only vaccine is to understand it and legislate against it
- **But there is still so much that we don't yet understand**

<sup>1</sup> <https://www.who.int/news-room/detail/02-05-2018-9-out-of-10-people-worldwide-breathe-polluted-air-but-more-countries-are-taking-action>

<sup>2</sup> Lelieveld, J. *et al.* The Contribution of Outdoor Air Pollution Sources to Premature Mortality on a Global Scale. *Nature* 2015, 525 (7569), 367–371

<sup>3</sup> <https://www.worldometers.info/coronavirus/coronavirus-death-toll/>

Time-series showing the global daily fatalities due to air pollution/ COVID-19

# Air pollution: an old problem



Air pollution episode	Date	Cause	Consequences
Meuse River Valley, Belgium	1930	High concentrations of SO <sub>2</sub> in the atmosphere during a temperature inversion	63 people died and thousand sick
Donora, Pennsylvania, USA	1948	High concentrations of SO <sub>2</sub> coupled with temperature inversion and foggy weather	20 people died due to cardiac respiratory disease and about half of town's 12,000 residents complained of cough, respiratory tract irritation, chest pain, headaches, nausea, and vomiting
Poza Rica, Mexico	1950	Natural gas plant inadvertently released hydrogen sulphide coupled with temperature inversion and foggy weather	22 people died and 320 were hospitalized
London, England	1952	A five day temperature inversion trapped deadly acid aerosols in the atmosphere	8000 deaths. Over 4000 succumbed to bronchitis, pneumonia, and respiratory and cardiac disease



# Air pollution: a brief history



*What we knew in the 1950s:*

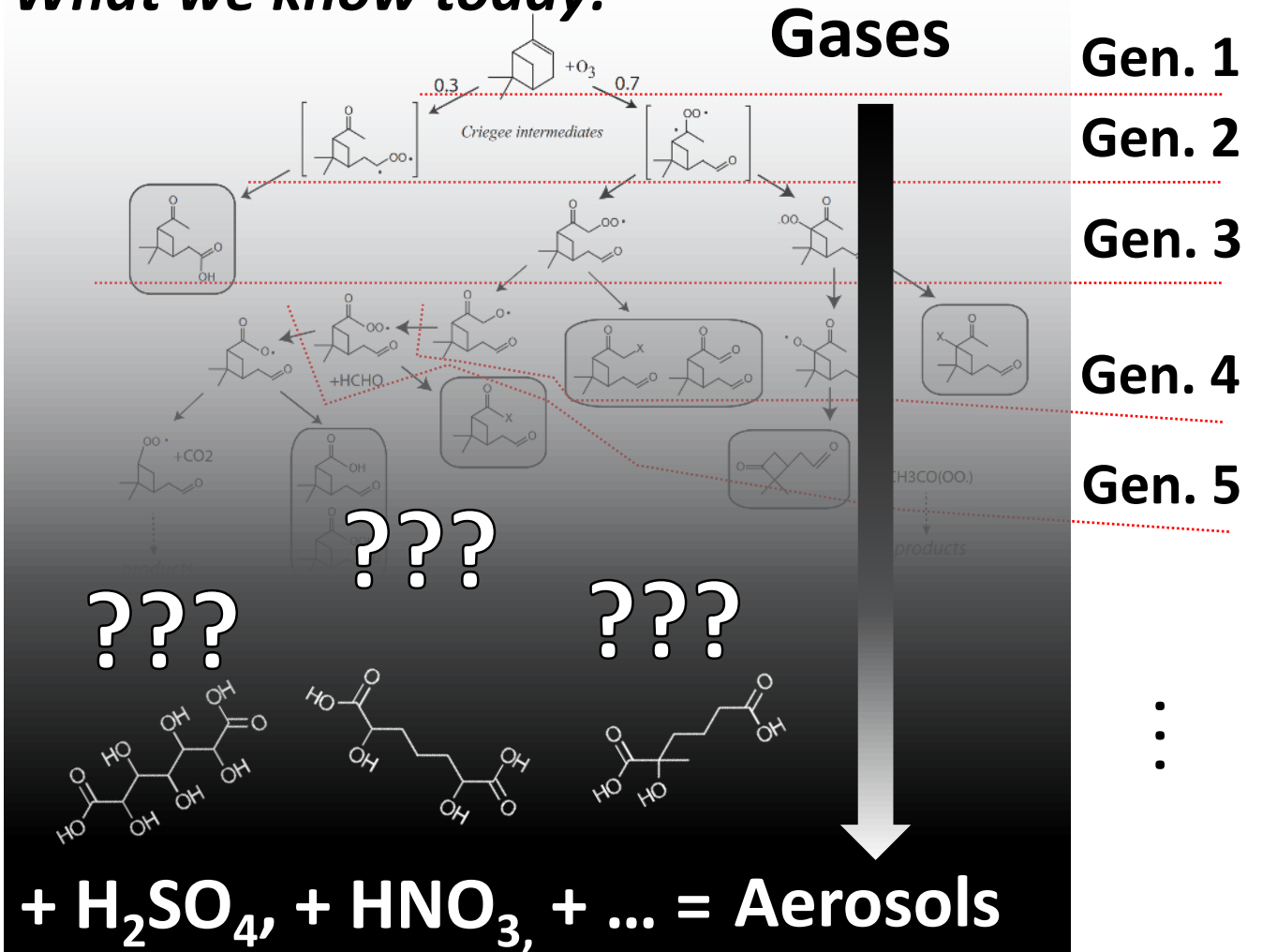
≡ **Hydrocarbons + NO<sub>x</sub>** ≡  
+ **sunlight + SO<sub>2</sub>** ≡

**peroxides**  
**aldehydes**  
**acids**  
???

**SMOG/Aerosols**

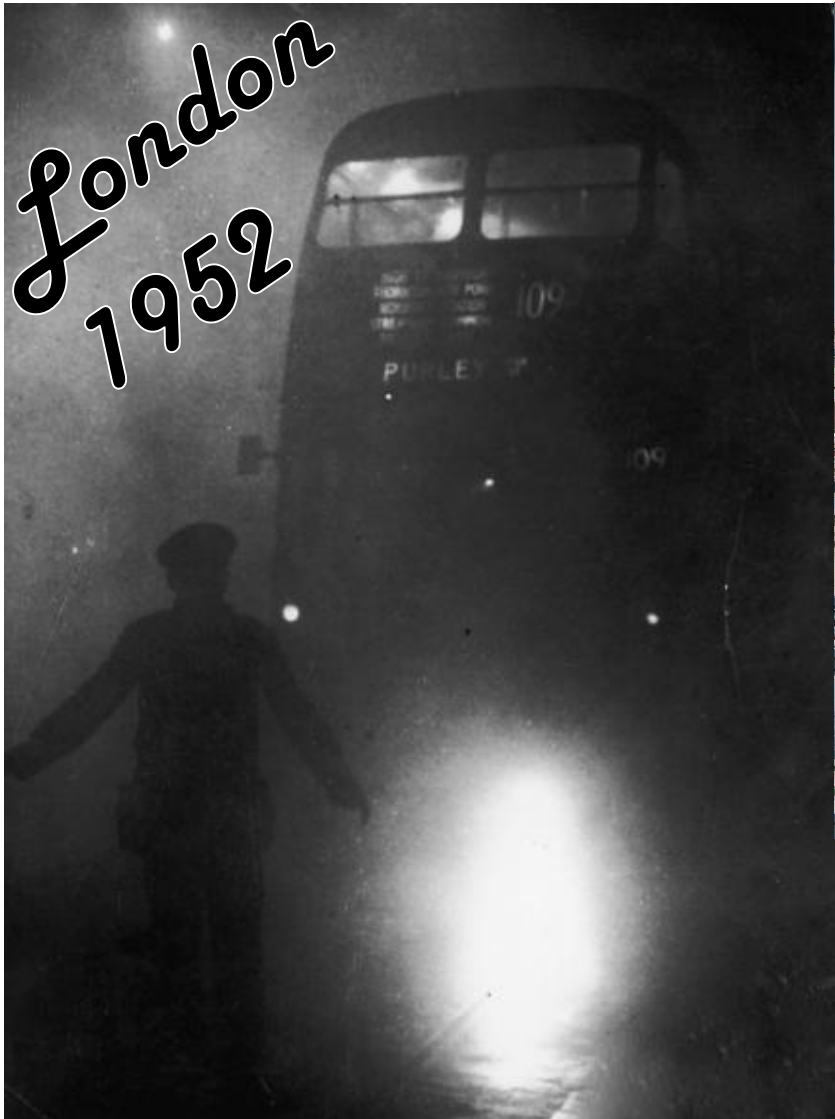
*Adapted (radically) from: Haagen-Smit (1952)*

*What we know today:*



*Sources: Valorso et al. (2011), Gao et al. (2004)*

# Is air pollution a thing of history?



**It depends on your perspective**



# Is air pollution a thing of history?



China, 2014



India, 2018



Nigeria, 2018



**Clearly not**

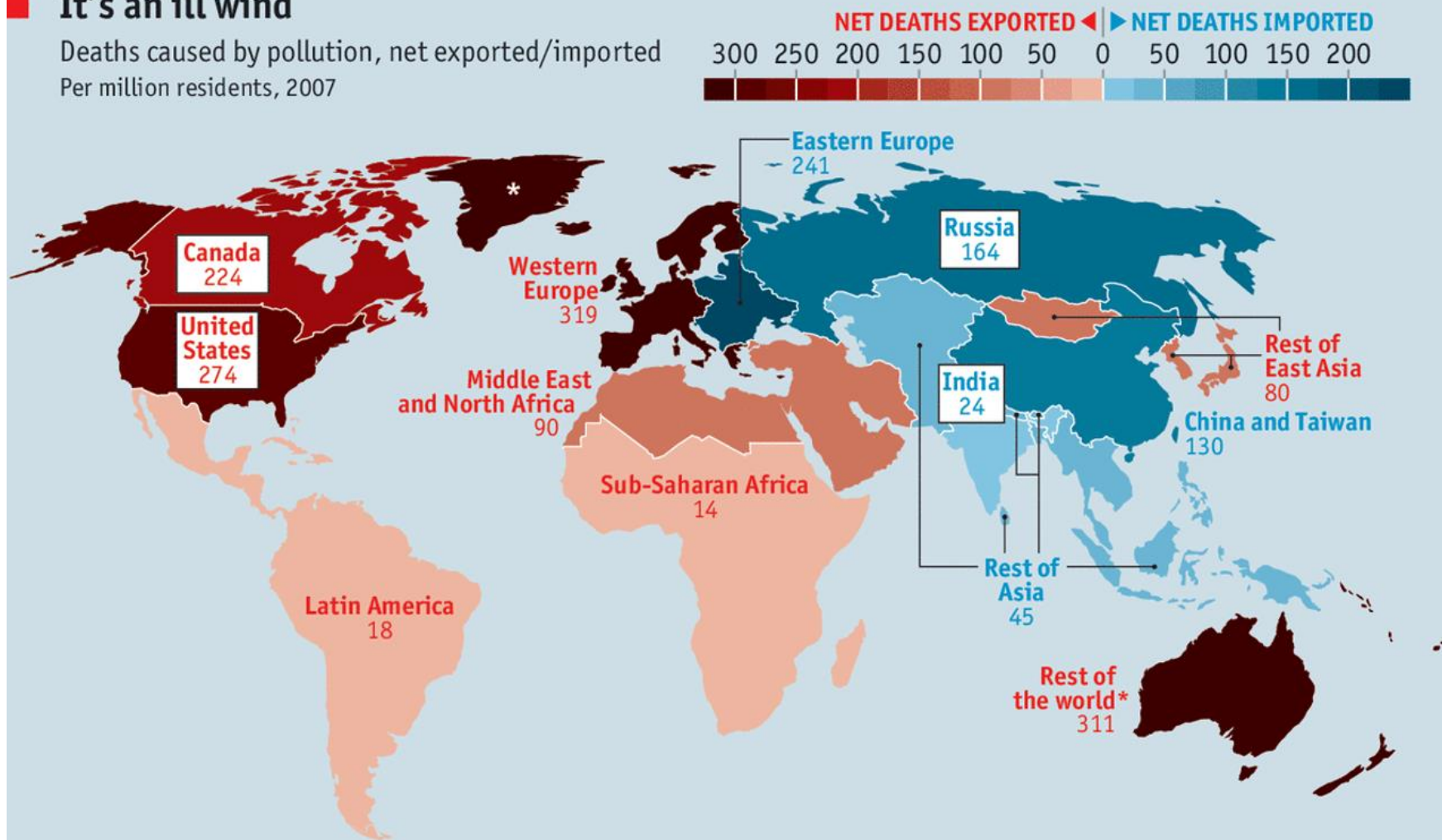


# The export of pollution



## It's an ill wind

Deaths caused by pollution, net exported/imported  
Per million residents, 2007

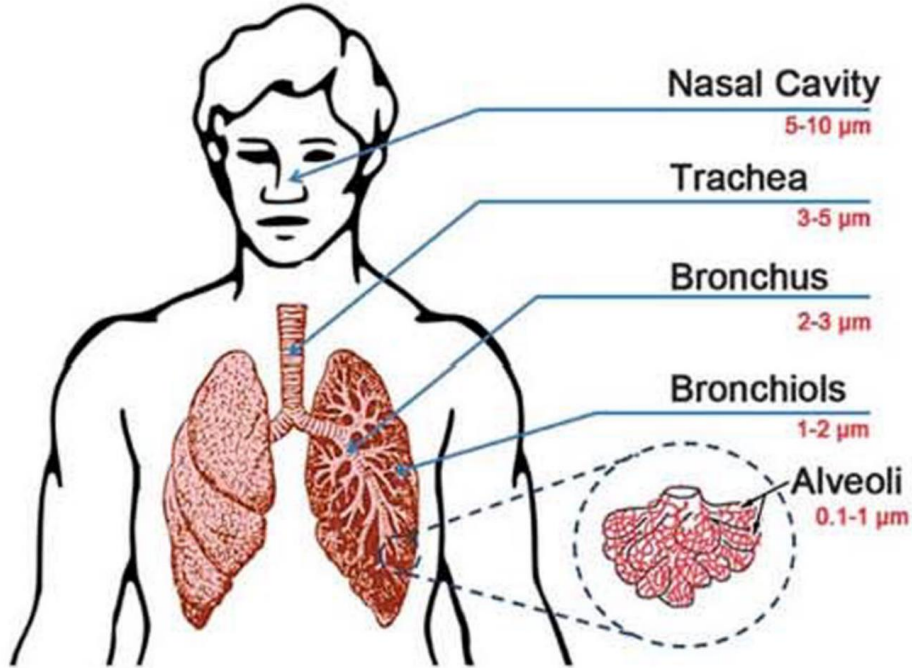


Sources: *Nature*; World Bank

\*"Rest of the world" includes Greenland

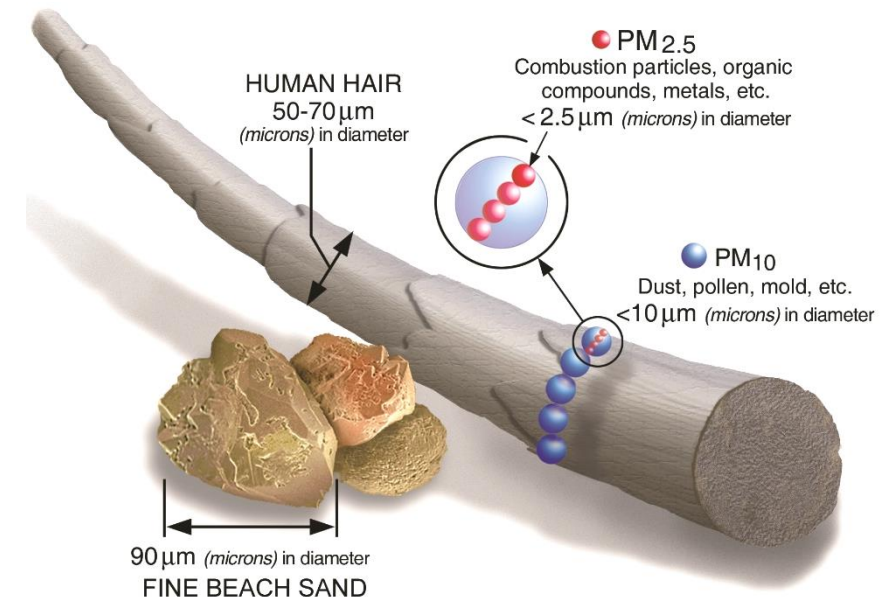
- The Global North is transitioning away from industrialization
- But our appetite for consumer goods and resources has increased
- As a consequence, we are seeing an explosion of pollution in the developing world
- Note: this figure is from 2007, the geopolitical picture is changing fast
- What about Africa?

# The health effects of air pollution



- Fine particulate matter is considered be the most dangerous component of air pollution
- These are the most aerodynamic/ mobile particles and they can reach the alveoli
- It is suspected that chemical composition may play a role in toxicity of particulates

- Lung cancer and chronic obstructive pulmonary disease are very bad outcomes
- But many other conditions are correlated with air pollution: asthma, dementia, COVID deaths
- Even behavioural patterns and cognitive ability

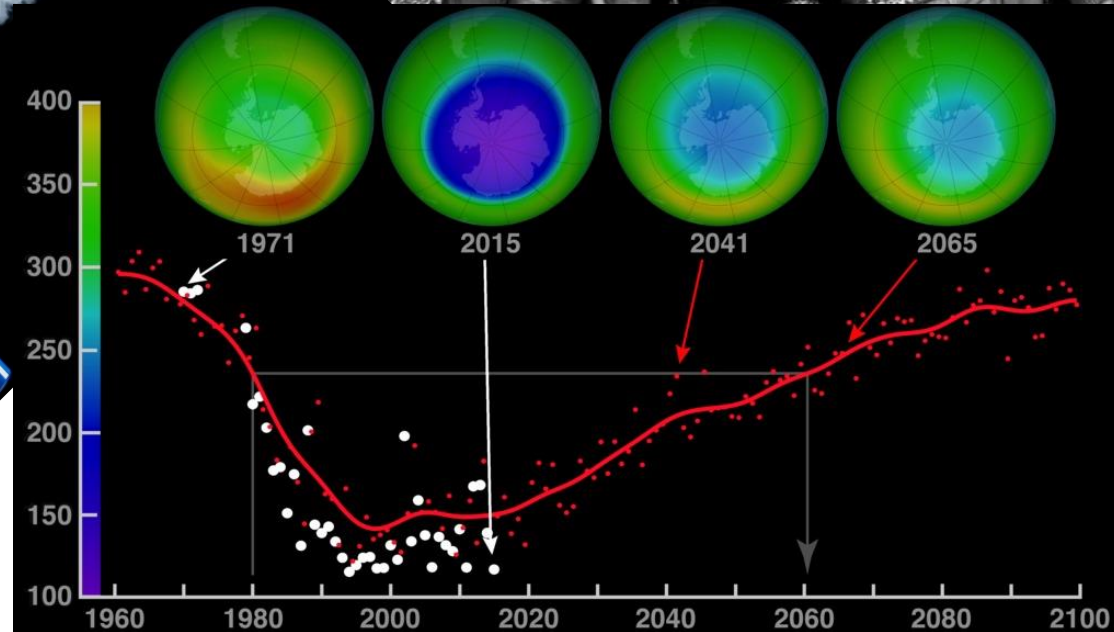




# The ozone hole

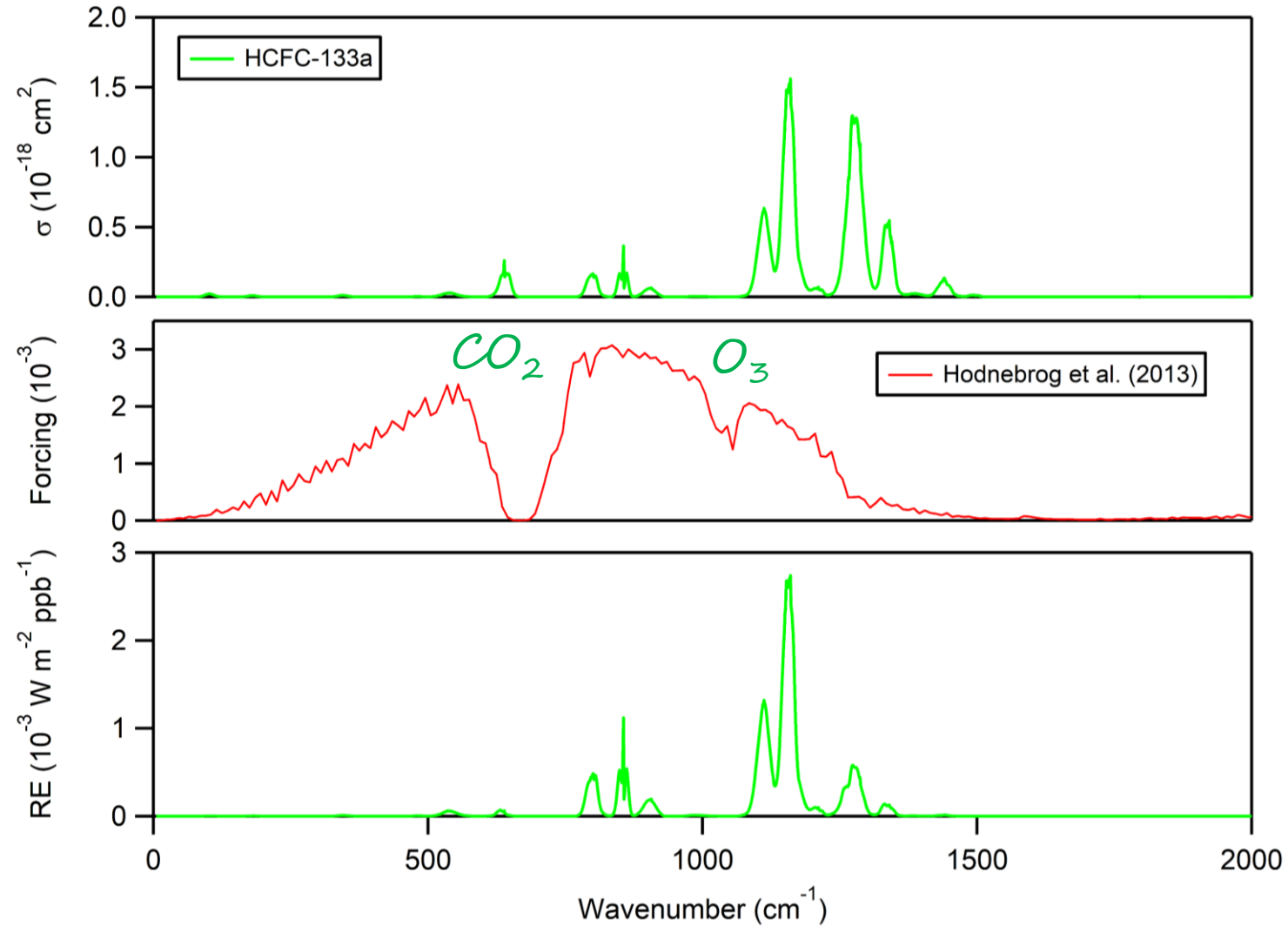


- The use of ozone-depleting substances starting in the 1960s has had a profound effect on the upper atmosphere
- These chemicals were used in a large variety of applications (not only hairspray!)
- This can be viewed as stratospheric air pollution



# Climatic effects

- Many pollutants (such as  $\text{CO}_2$ ) absorb infrared radiation
- Some pollutants absorb strongly in the atmospheric “window”  $\rightarrow$  radiative efficiency (RE)
- Some pollutants are highly persistent in the atmosphere, leading to a high global warming potential (GWP)
- Some of the worst greenhouse gases can have a GWP 1000s of times higher than  $\text{CO}_2$



# Conclusions:

## Part 1 – Why does air pollution matter?

- Air pollution is estimated to kill millions of people each year
- It impacts the lives of billions
- It's less problematic now than it was (in the developed world)
- But it's becoming worse (in the developing world)
- Air pollution now kills more people than tobacco (a lifestyle choice)
- Breathing is not a lifestyle choice, it is a necessity!
- + we should not forget the climatic effects that will impact the whole world

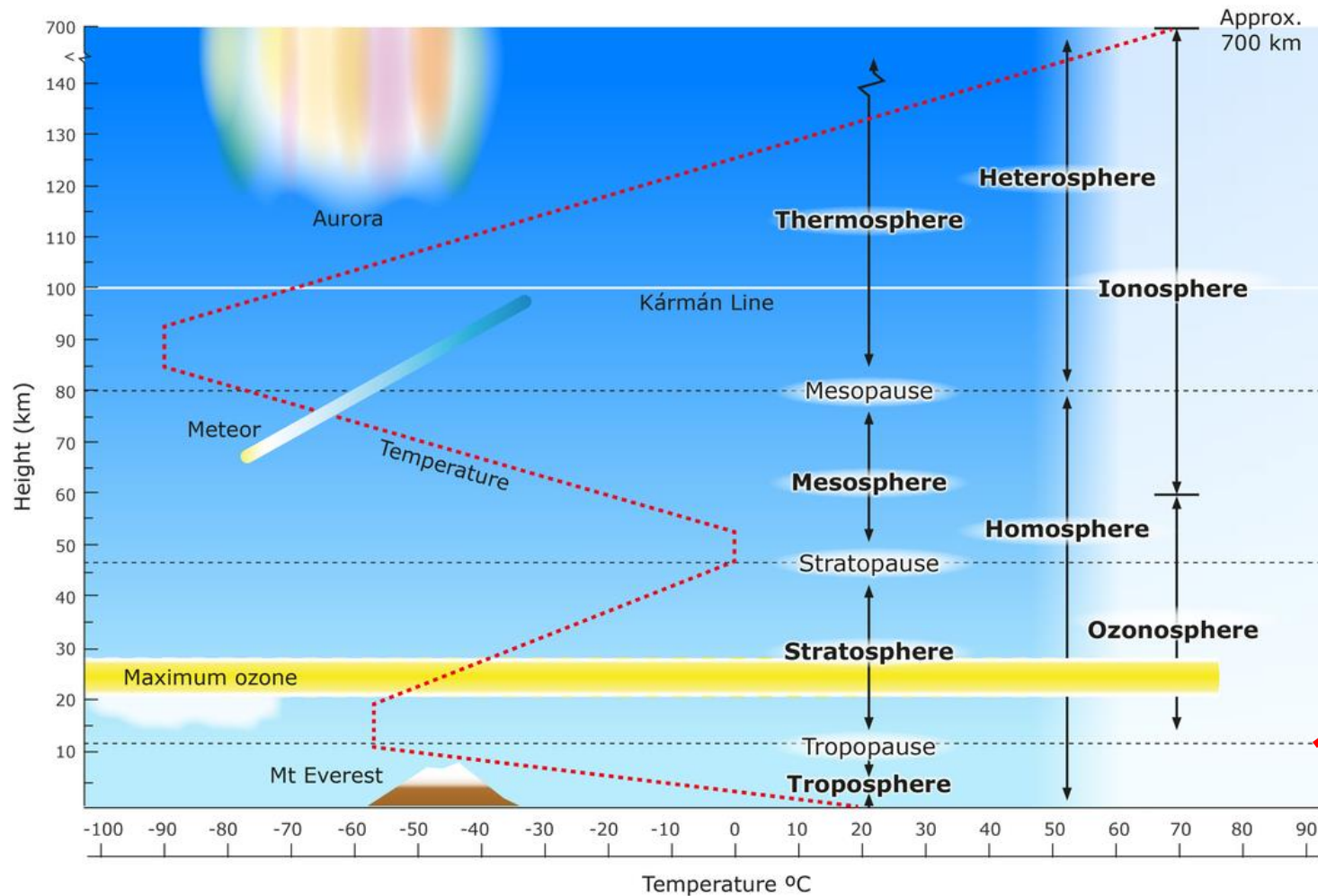




# Part 2 – Getting to know the atmosphere



# Physical attributes (structure)



- The atmosphere is categorized as a set of layers
- We spend almost all of our time in the planetary boundary layer

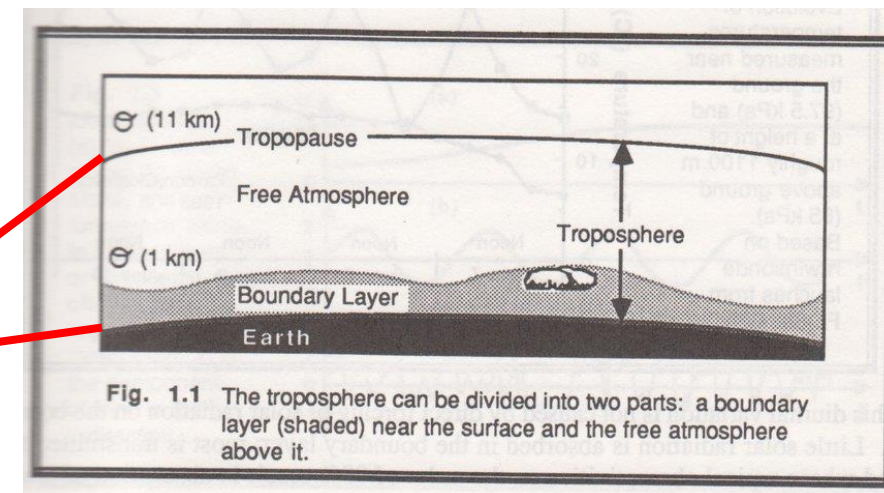


Fig. 1.1 The troposphere can be divided into two parts: a boundary layer (shaded) near the surface and the free atmosphere above it.

# Physical attributes (atmospheric pressure)

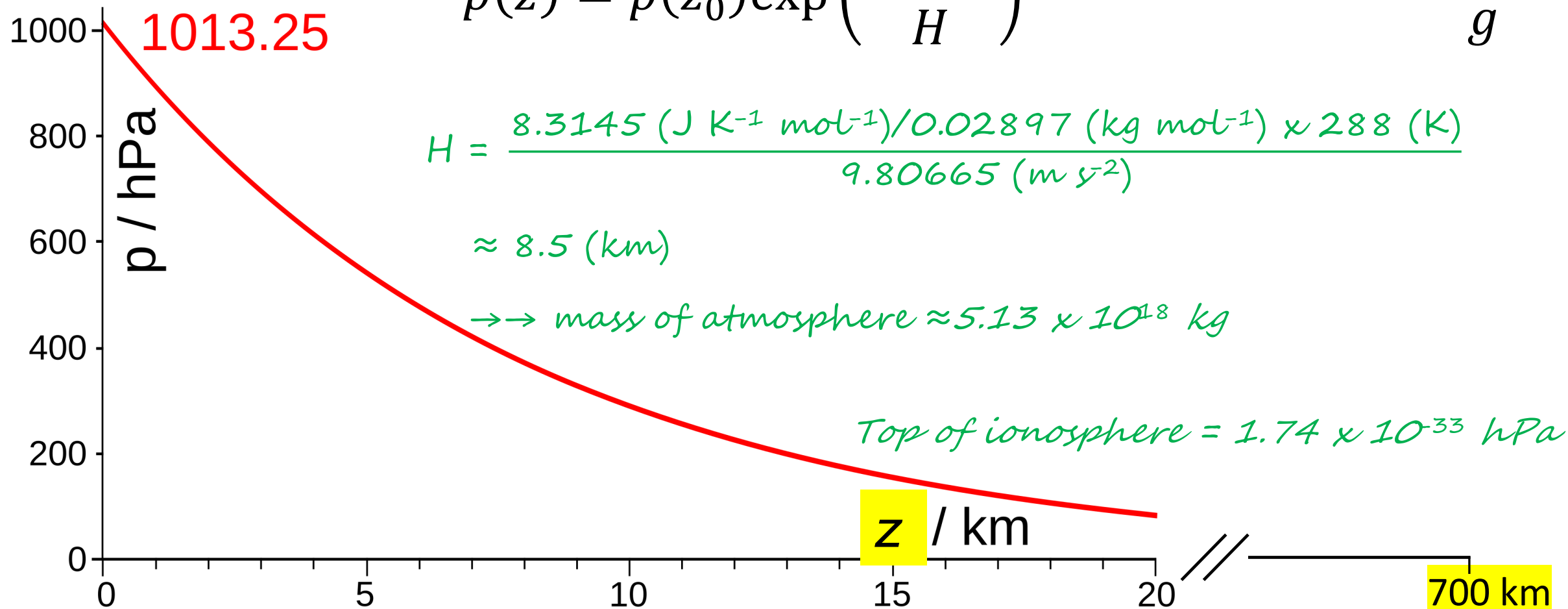
$$p(z) = p(z_0) \exp\left(\frac{z - z_0}{H}\right) \quad \text{where:} \quad H = \frac{RT}{g}$$

$$H = \frac{8.3145 \text{ (J K}^{-1} \text{ mol}^{-1}) / 0.02897 \text{ (kg mol}^{-1}) \times 288 \text{ (K)}}{9.80665 \text{ (m s}^{-2})}$$

$$\approx 8.5 \text{ (km)}$$

$$\rightarrow \rightarrow \text{mass of atmosphere} \approx 5.13 \times 10^{18} \text{ kg}$$

$$\text{Top of ionosphere} = 1.74 \times 10^{-33} \text{ hPa}$$



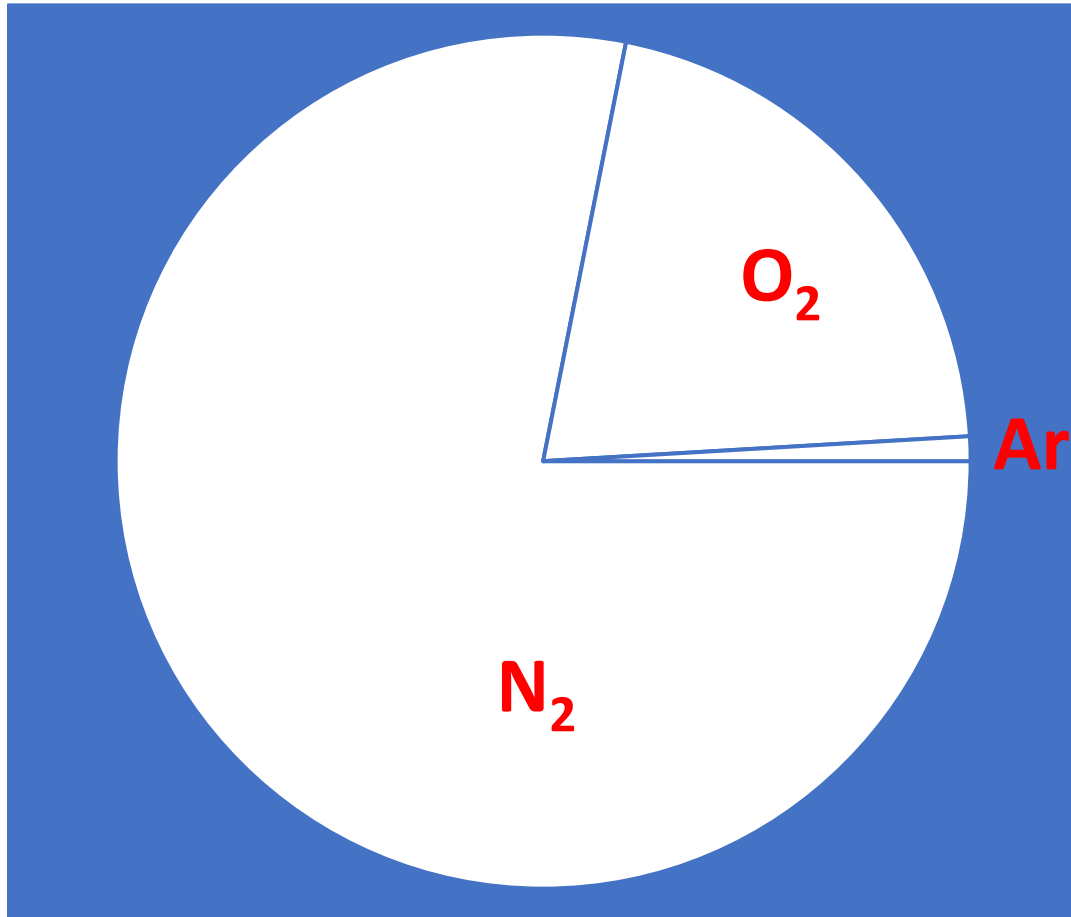
# Physical attributes (in context)



- In terms of spatial extent: the atmosphere is to the solid Earth, what an apple's skin is to its diameter
- In terms of mass: it is a single eyelash lying next to the apple...
- From a physical point of view, the atmosphere is therefore a thin, insubstantial layer



# Chemical composition (major components)



$$N_2 = 78.08\%$$

$$O_2 = 20.95\%$$

$$Ar = 0.93\%$$

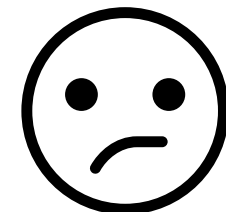
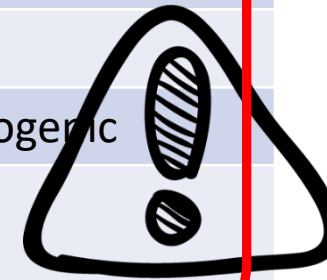
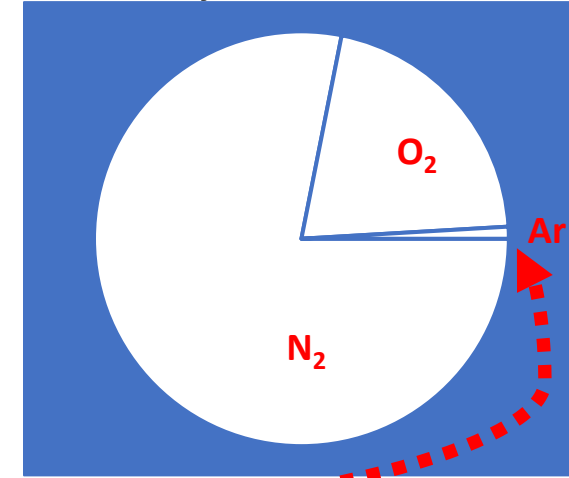
$$99.96\%$$

Most of the  
atmosphere is very  
boring!



# Chemical composition (trace components)

Gas	Mixing ratio	Source(s)
Carbon dioxide	420 (ppm)	Combustion, ocean, biosphere
Methane	1.7 (ppm)	Biogenic, anthropogenic
Helium	5.2 (ppm)	Radioactive decay, anthropogenic
Nitrous oxide	0.31 (ppm)	Biogenic, anthropogenic
Ozone (stratosphere)	0.5–10 (ppm)	Photochemical
Ozone (troposphere)	10–500 (ppb)	Photochemical
Carbon monoxide	50–200 (ppb)	Photochemical, combustion
Formaldehyde	0.1–1 (ppb)	Photochemical, anthropogenic
NMHCs	ppqs–ppbs	Photochemical, anthropogenic, biogenic
Hydroxyl radical (daytime)	~ 40 (ppq)	Photochemical



Total = 0.0438% (+99.96%) = 100.0038%

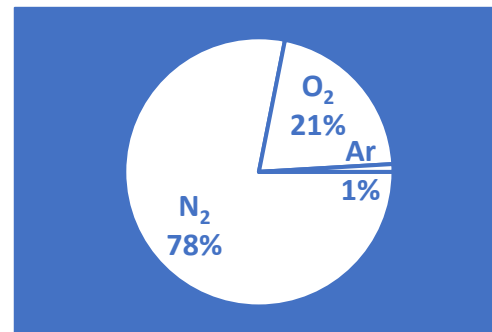
# How has Earth's atmosphere changed?

- Earth's atmosphere was very different millions/billions of years ago
- The evolution of photosynthetic organisms transformed much of the CO<sub>2</sub> into O<sub>2</sub>
- Further back in time, Earth's atmosphere would have contained much more of the most abundant components of the solar system: hydrogen and helium
- Because of atmospheric escape, most has been lost to space

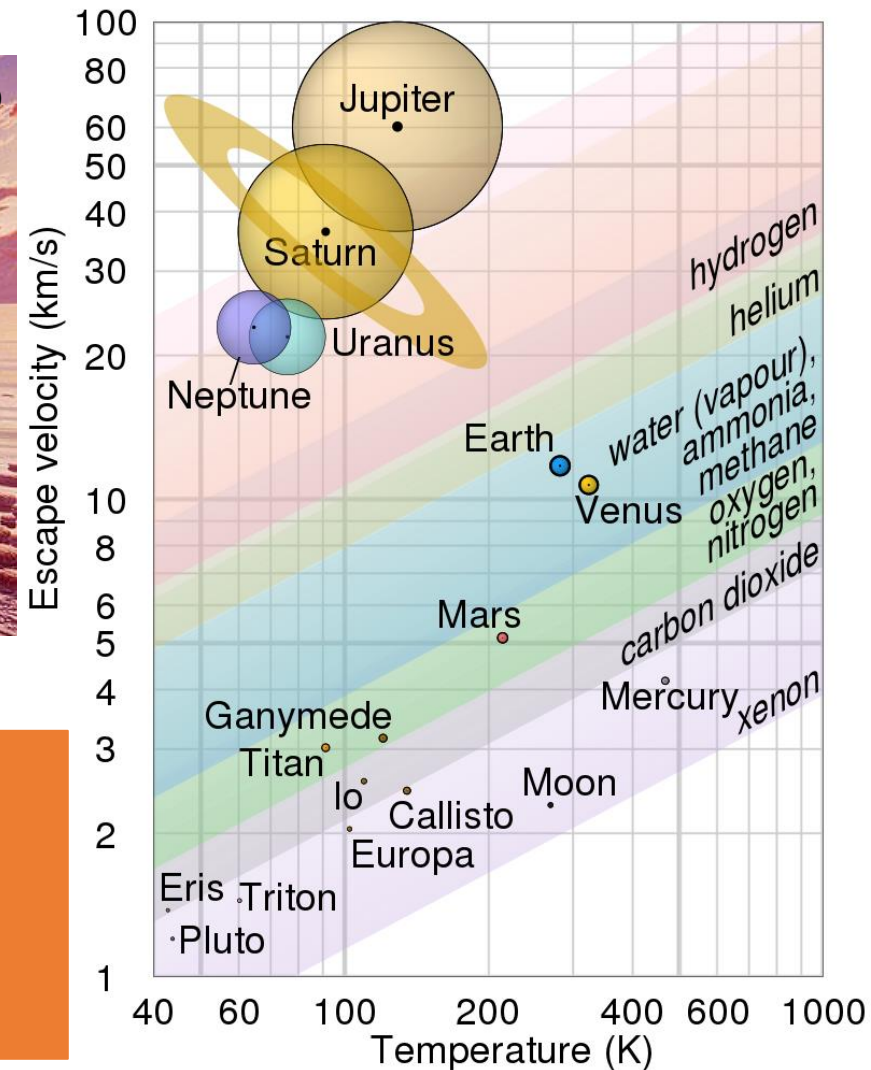
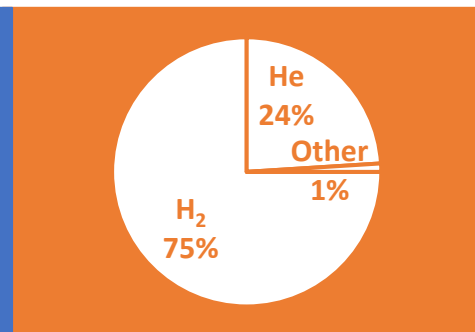
Prebiotic Earth



Earth

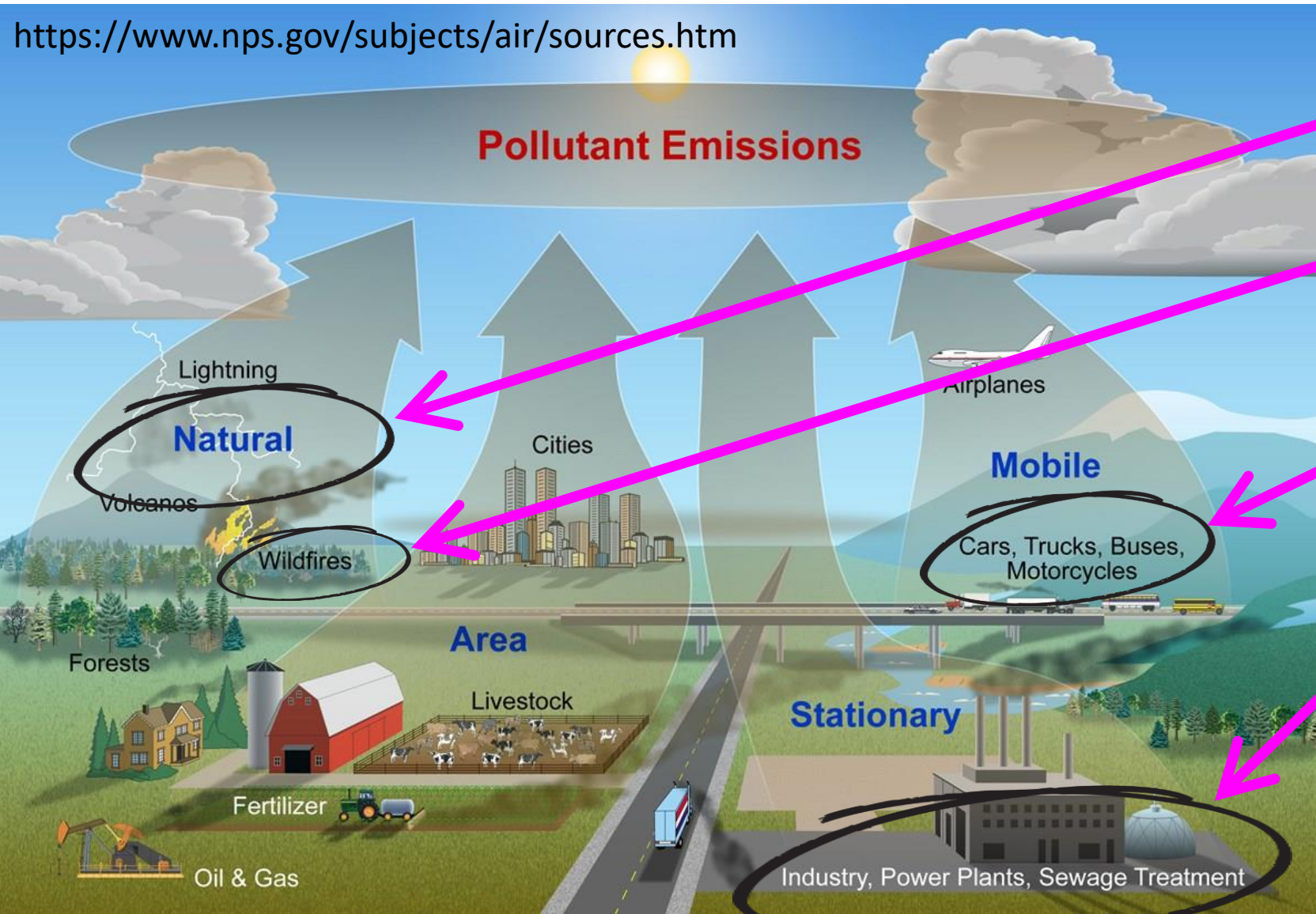


Jupiter



# Atmospheric emissions

<https://www.nps.gov/subjects/air/sources.htm>



- 55,000 different terpenoid species known to science
- Wildfires, >10,000 species identified in the tar-ball component of a single burning plume
- Hydrocarbon fuels are complex mixtures, containing  $10^3$ – $10^4$  species (maybe more)
- >75,000 industrial chemicals registered in environmental monitoring networks (US EPA CPDAT)



# A paradox: low concentration, big effect

- How can things at such low concentrations be so important?

- **Possibility 1:** the mass flux is high

*sources* → **[isoprene]** → *sinks*

*Field concentration: ppt*

*Atmospheric emission:  
 $5-7.5 \times 10^6 \text{ tons yr}^{-1}$*

- **Possibility 2:** a molecule has a catalytic effect

*e.g. stratospheric chlorine sources (see later)*

- **Possibility 3:** high toxicity

*high specific toxicity / high collective toxicity*

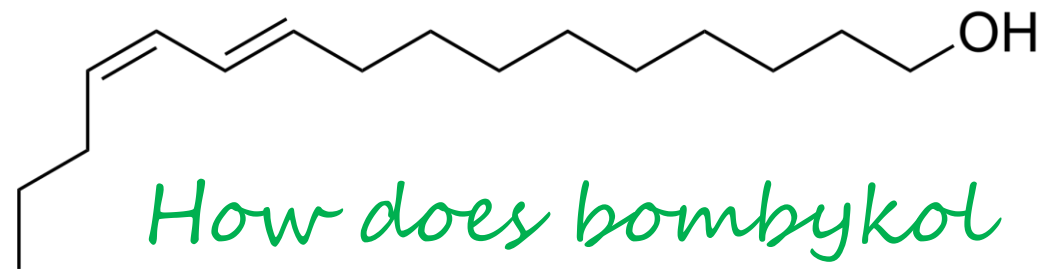
- **Possibility 4:** extremely sensitive effect

*aerosol nucleation / infrared absorption /  
catalytic behaviour / biological role*

# VOCs and ecosystems



- The most sensitive VOC detectors on Earth are not in our laboratories
- Certain species of insect are hypothesized to be able to detect single molecules
- We are capable of measuring 10s of ppq levels with certain mass spectrometric and fluorescence techniques (if we are lucky)
- For a sample volume of 1 cm<sup>3</sup> at atmospheric pressure and room temperature, this is  $>2.5 \times 10^5$  molecules



*How does bombykol behave in the atmosphere?*



# Conclusions

## Part 2 – Getting to know the atmosphere

- We live in a thin, fragile and insubstantial part of the Earth system
- Most of the atmosphere is relatively inert
- Earth's atmosphere has changed a lot since it was formed
- Trace components affect air quality, climate, ecosystems
- These are the things that react...



# Part 3 – Chemical reactions in the troposphere



# The atmosphere as a combustion engine



- Much like an engine, the atmosphere is an oxidizing environment
- The volatile organic compounds (VOCs) emitted to the atmosphere represent the fuel
- What we learned at school:



- However, under ambient (and most other) conditions, these things don't react together...

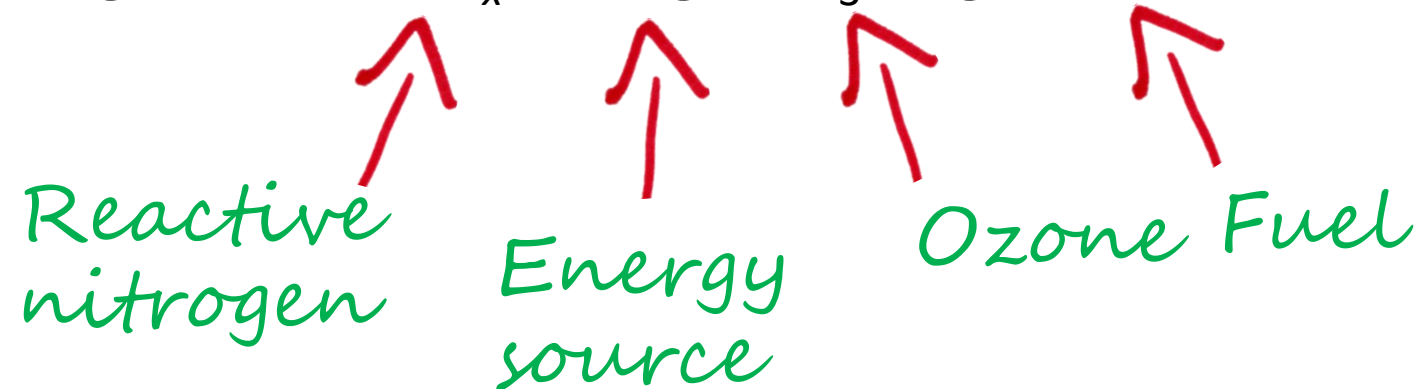




# Getting things started...



- Atmospheric chemistry is a complex set of cycles, and it's difficult to choose a starting point!
- Remember, the atmosphere is mostly  $N_2$ ,  $O_2$  and Ar with trace VOCs
- But none of these things react together...
- We need some extra ingredients:  $NO_x$ , sunlight,  $O_3$ , organic molecules

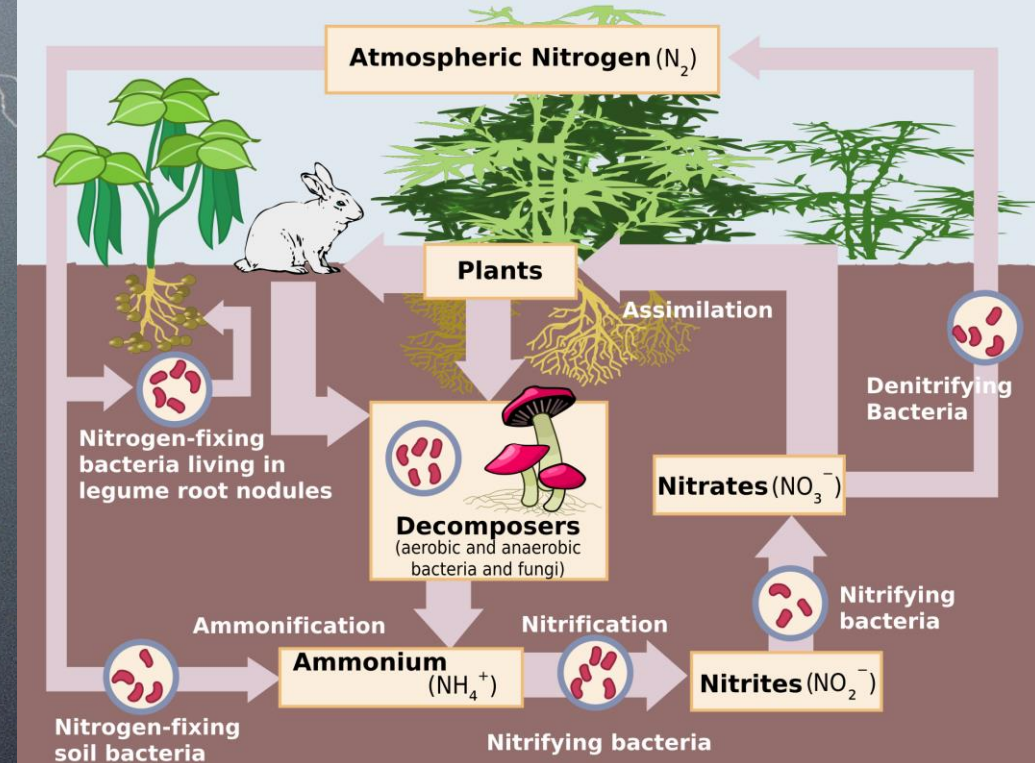


- These create the conditions for the autocatalytic oxidizing environment of the troposphere



# Ingredient 1: fixed nitrogen

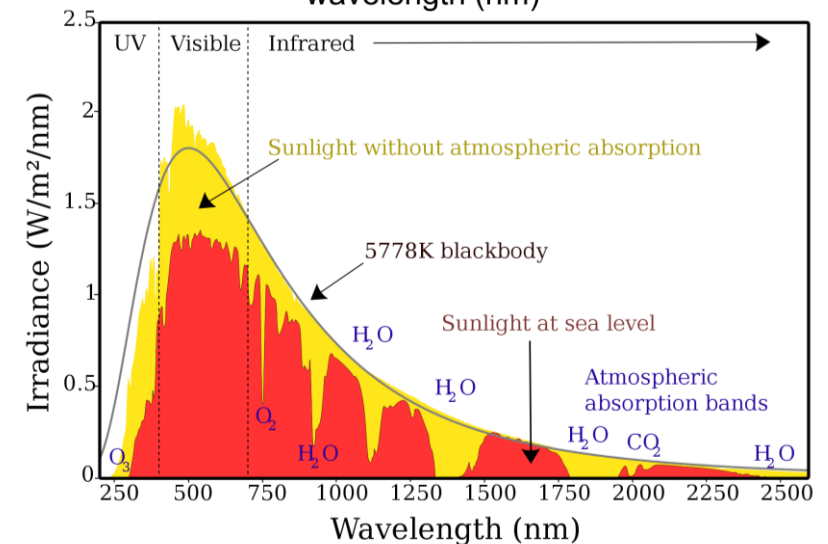
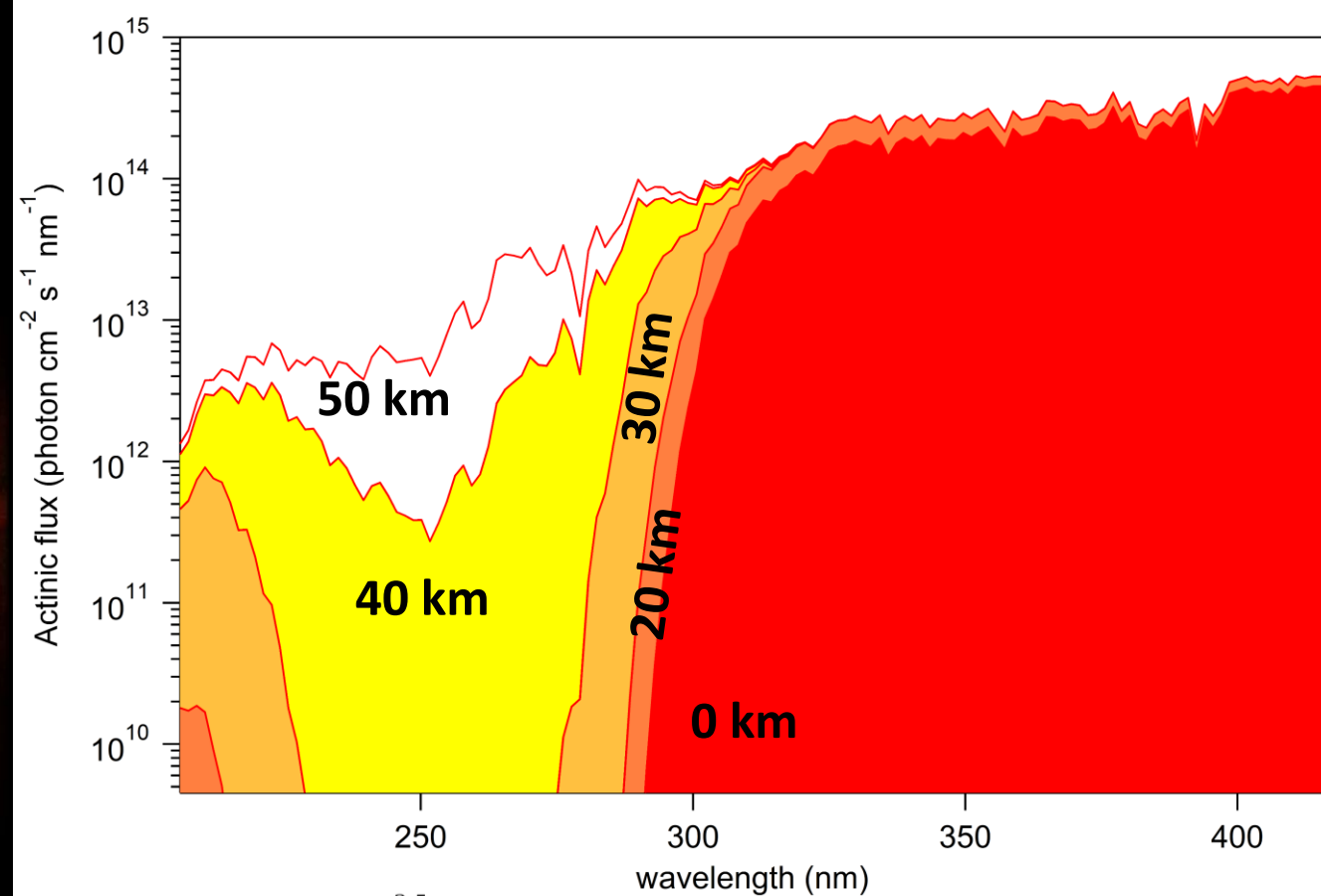
- Most nitrogen is present in a highly unreactive form:  $\text{N}\equiv\text{N}$  (BDE =  $226 \text{ kcal mol}^{-1}$ )
- Nitrogen “fixation” may be biotic or abiotic
- Biotic: biochemical reactions involving the nitrogenase enzyme
- Abiotic: generally involves extreme heat (the core of lightning bolt/ a high-temperature combustion engine)





# Ingredient 2: sunlight

- The sun is the energy source
- Top of the atmosphere: solar irradiance resembles 5778K black body
- Ground-level: the layers above act as a filter
- If a molecule absorbs enough energy it can transform...



# Photodissociation

- Reaction  $AB + h\nu \rightarrow A + B$
- Rate of reaction is given by:

$$-\frac{d[AB]}{dt} = J \times [AB]$$

- $J$ , the photolysis frequency, is described as follows:

$$J = \int_{\lambda_1}^{\lambda_2} F(\lambda) \times \Phi(\lambda) \times \sigma(\lambda)$$

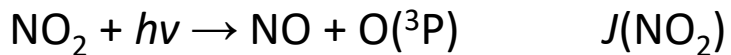
Where:  $F(\lambda)$  = actinic flux

$\sigma(\lambda)$  = absorption cross section

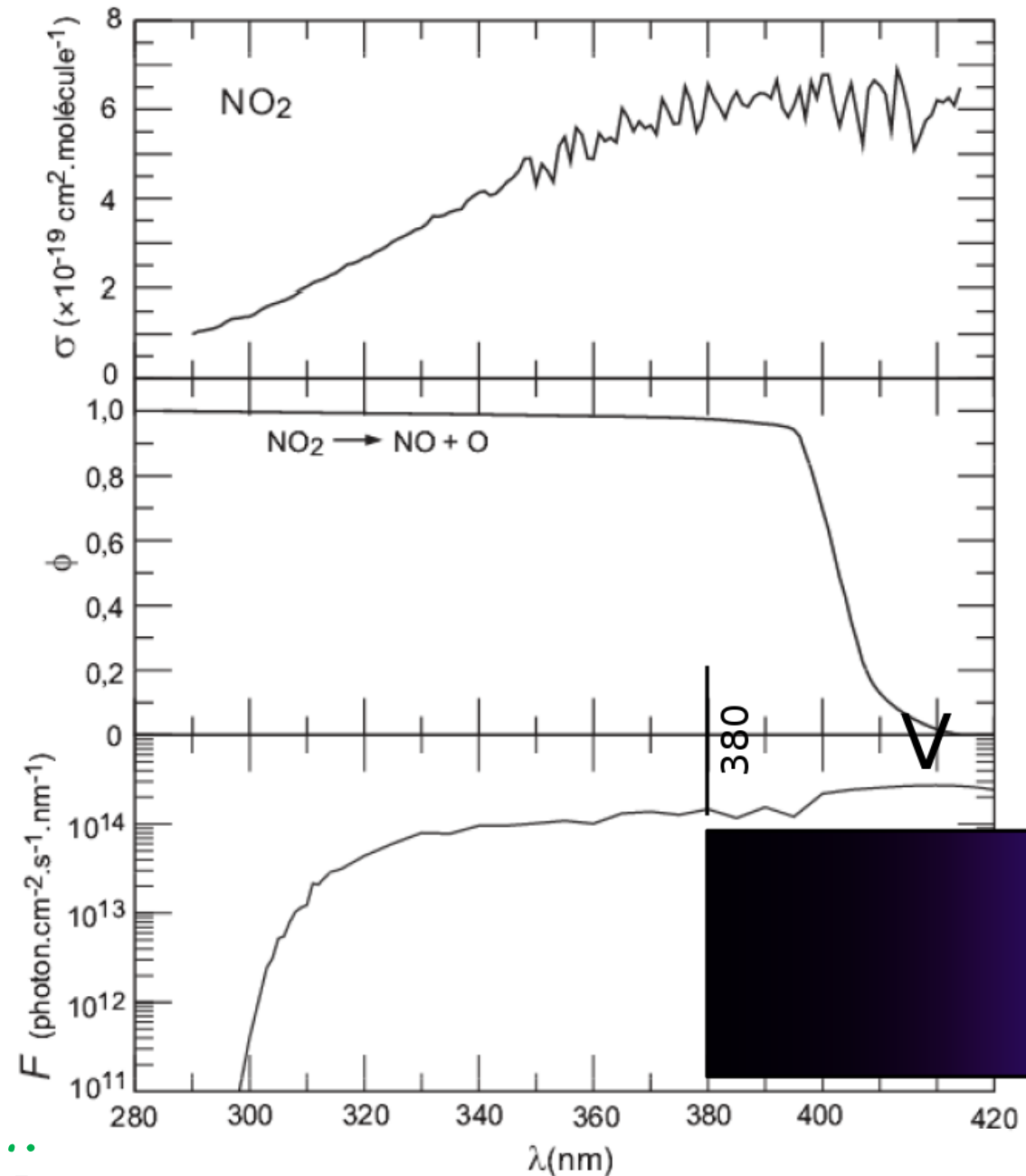
$\Phi(\lambda)$  = quantum yield

$$\Phi(\lambda) = \frac{\text{\#photons leading to reaction}}{\text{\#photons absorbed}}$$

- Example:

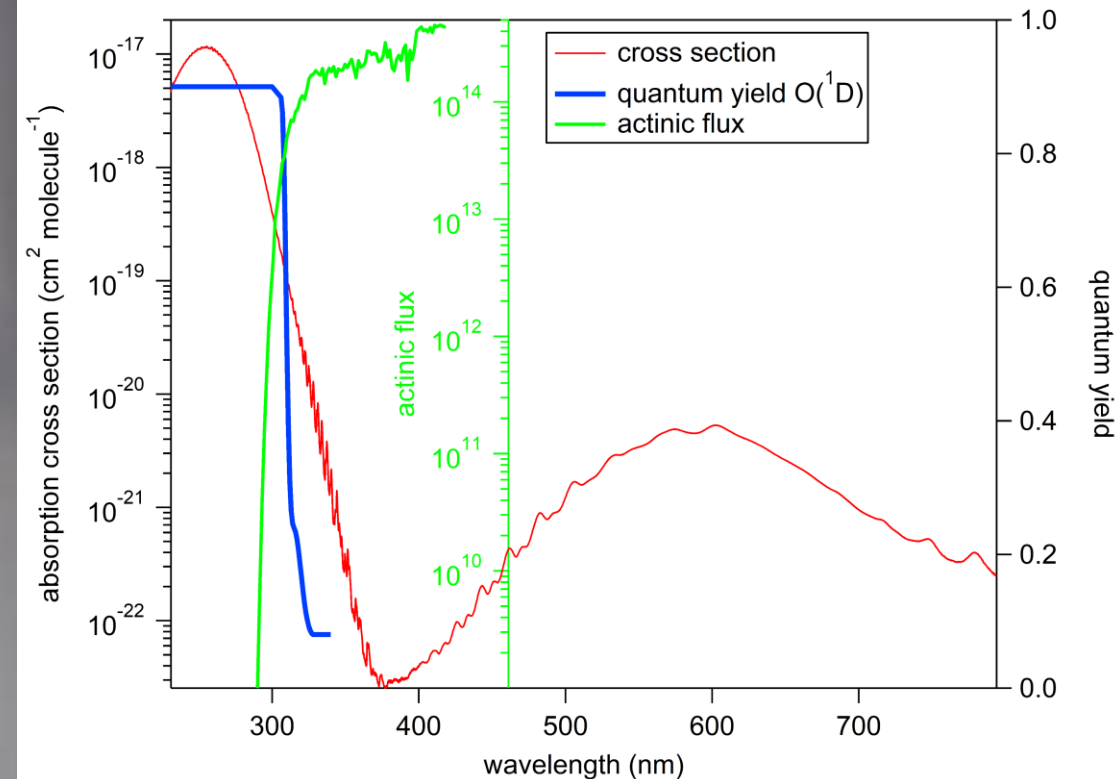
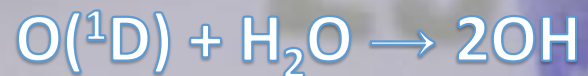
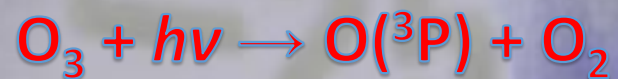
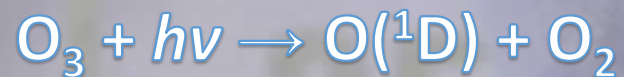


*However, there are other processes besides dissociation...*



# Ingredient 3: ozone

- Now we have some fixed nitrogen and some sunlight:



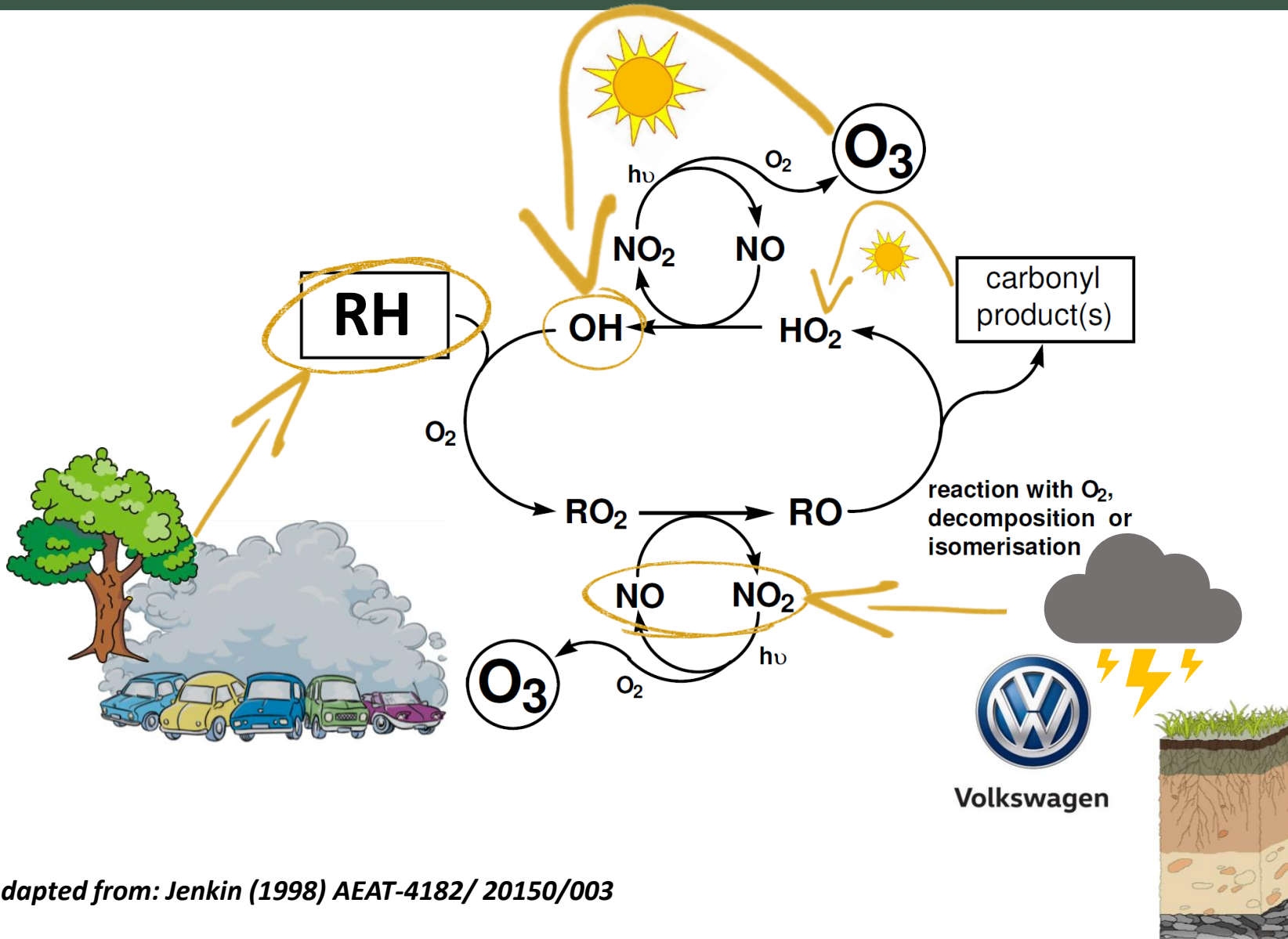


# Ingredient 4: fuel

- Finally, we need the fuel
- In principle, this can be any VOC
- Some fuels are more reactive than others
- This leads to differences in local and regional effects
- Some fuels are “cleaner” than others
- i.e. contain more carbon → longer chemical chain length
- → more efficient aerosol precursors

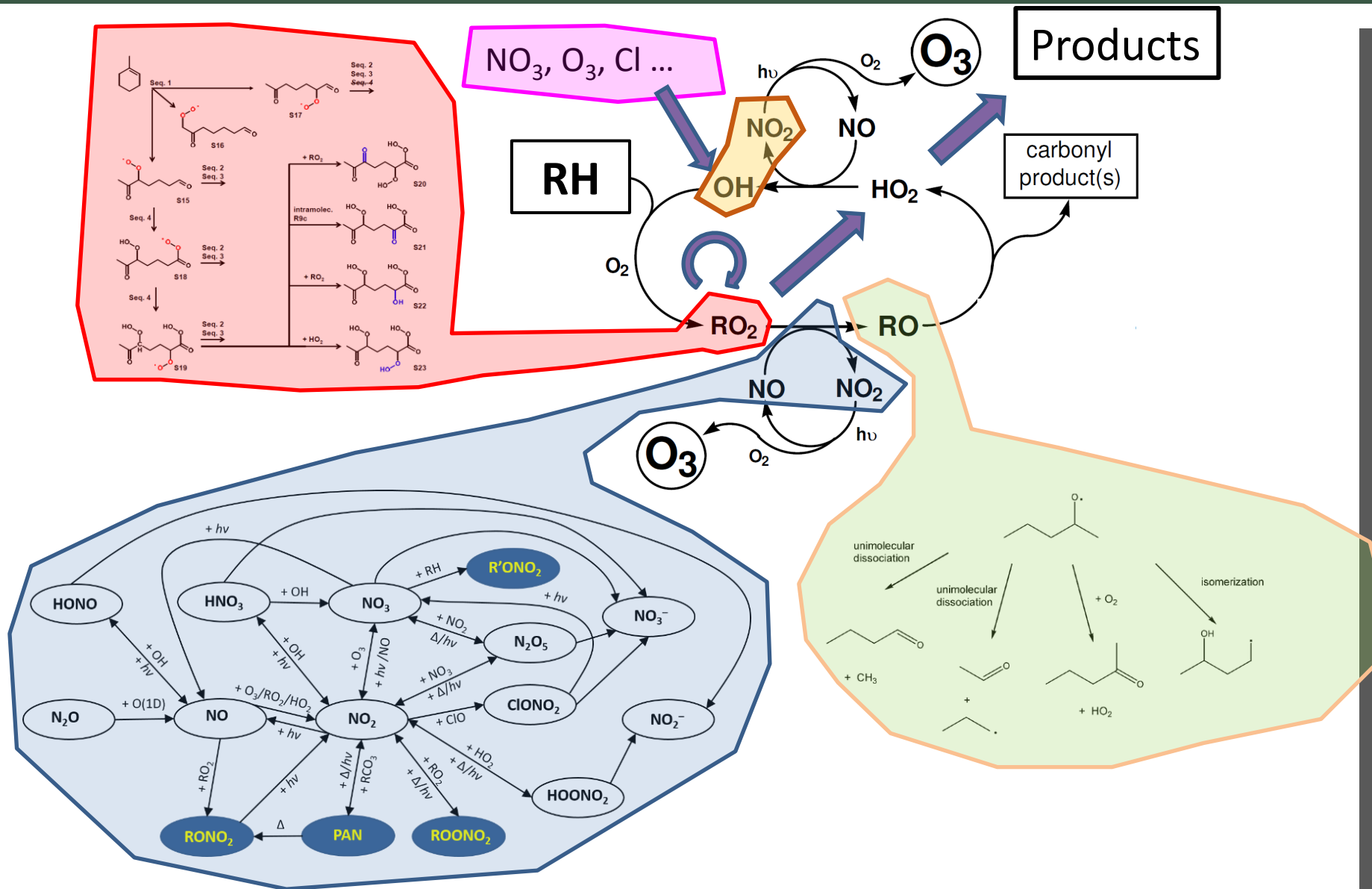


# Oxidation in the troposphere



- VOCs are emitted from many sources: trees, vehicles, industries, etc.
- They are consumed by OH, but other oxidants are important (NO<sub>3</sub>, O<sub>3</sub>, Cl)
- **Oxidants are produced in a collectively autocatalytic set**
- Here, VOCs are transformed into a sequence of carbonyls of diminishing size → CO<sub>2</sub>  
However...

# But life is never that easy...



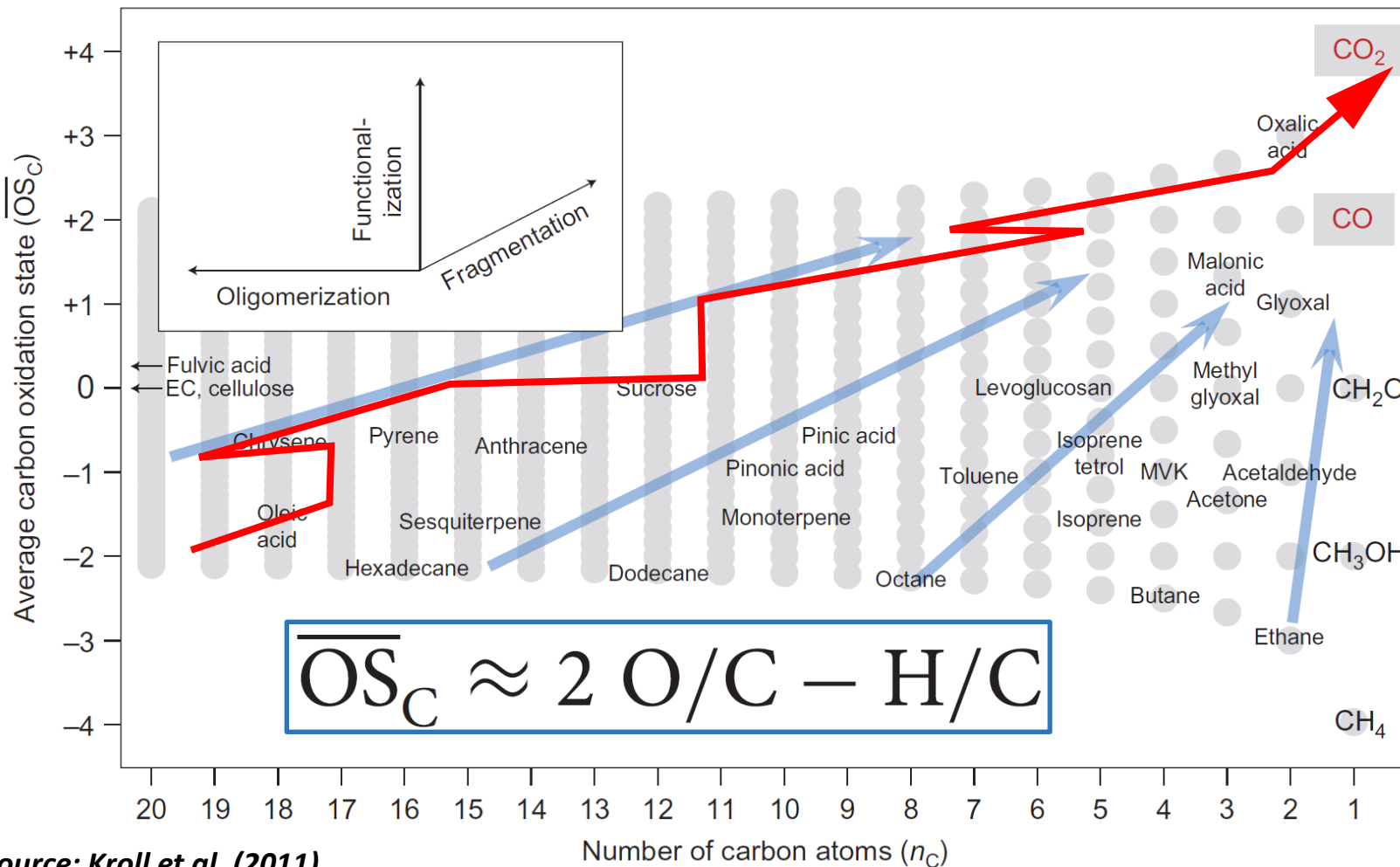
• There are many branching channels to consider

•  $\text{RO}_2$  and  $\text{RO}$  have a high capacity for isomerization and fragmentation

• Separate chemical cycles involving nitrogen and sulphur are important

• What if R contains a heteroatom?

# Meandering through oxidation space



Source: Kroll et al. (2011)

- Atmospheric chemistry is a very cold, very slow combustion
- Eventually, a “fuel” may be converted to CO<sub>2</sub> and H<sub>2</sub>O
- **But the path through this space is not simple or direct**
- At what point is atmospheric chemistry unknowable/unpredictable?



# The hydroxyl radical

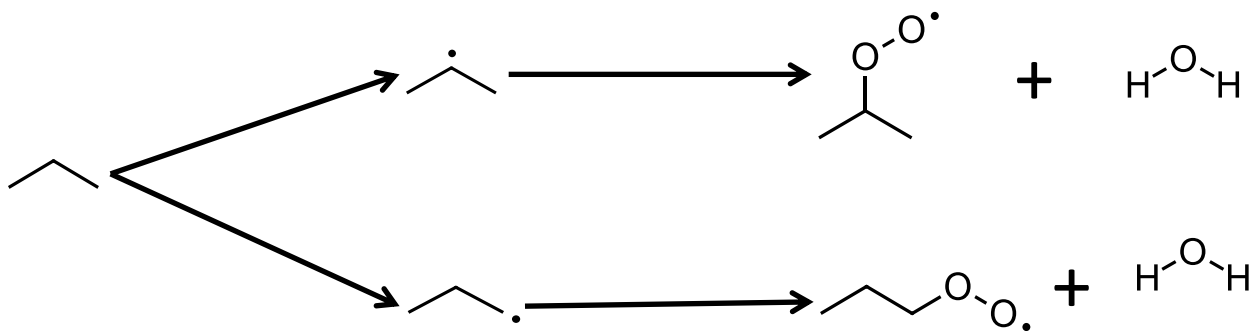


- Also known as the “detergent of the atmosphere”
- Daytime sources:  $\text{O}_3 + hv$ ,  $\text{HONO} + hv$ ,  $\text{RCHO} + hv$
- Nighttime source (minor): alkene ozonolysis
- Sinks: reactions with VOCs, reaction with  $\text{NO}_2$
- Ambient concentration:  $\sim 1 \times 10^6$  molecule  $\text{cm}^{-3}$
- OH has the balance between high reactivity and high enough concentration to be an effective cleaner of the atmosphere

# Hydroxyl radical reactions



- Hydrogen abstraction

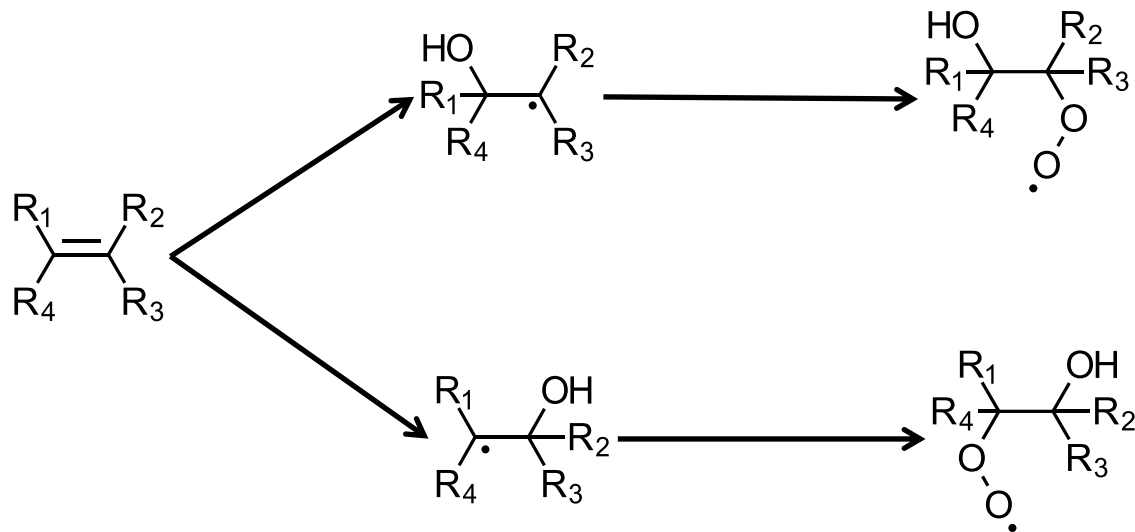


- Comparatively high reactivity with alkanes and other unsaturated species

# Hydroxyl radical reactions



- Hydroxyl addition



- High reactivity towards alkenes and aromatics
- Add alcohol functionality to alkenes

# The nitrate radical



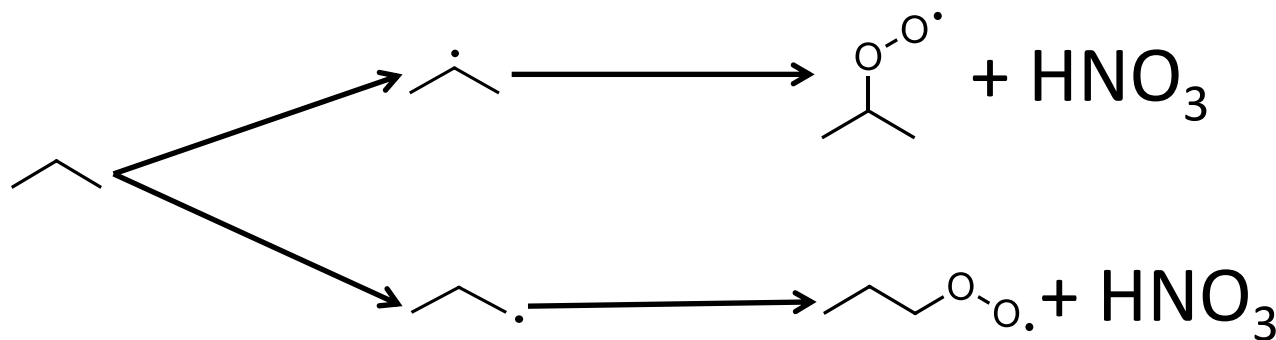
- The nitrate radical can be the most effective oxidant at night
- Nitrate is part of the following equilibrium:  $\text{NO}_2 + \text{NO}_3 \rightleftharpoons \text{N}_2\text{O}_5$
- $\text{NO}_3$  is most important in areas that are affected by anthropogenic  $\text{NO}_x$  and ozone
- Sinks: photolysis, reactions with alkenes and NO
- Ambient concentration:  $\sim 5 \times 10^8$  molecule  $\text{cm}^{-3}$
- $\text{NO}_3$  has a very high photolysis cross section in the actinic region, so it is effectively lost during the day



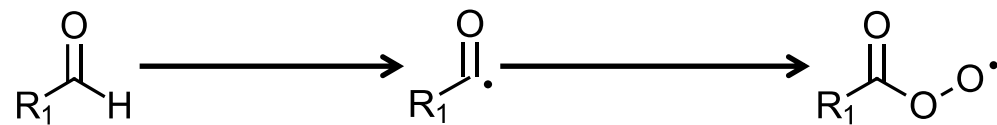
# Nitrate radical reactions



- Hydrogen abstraction



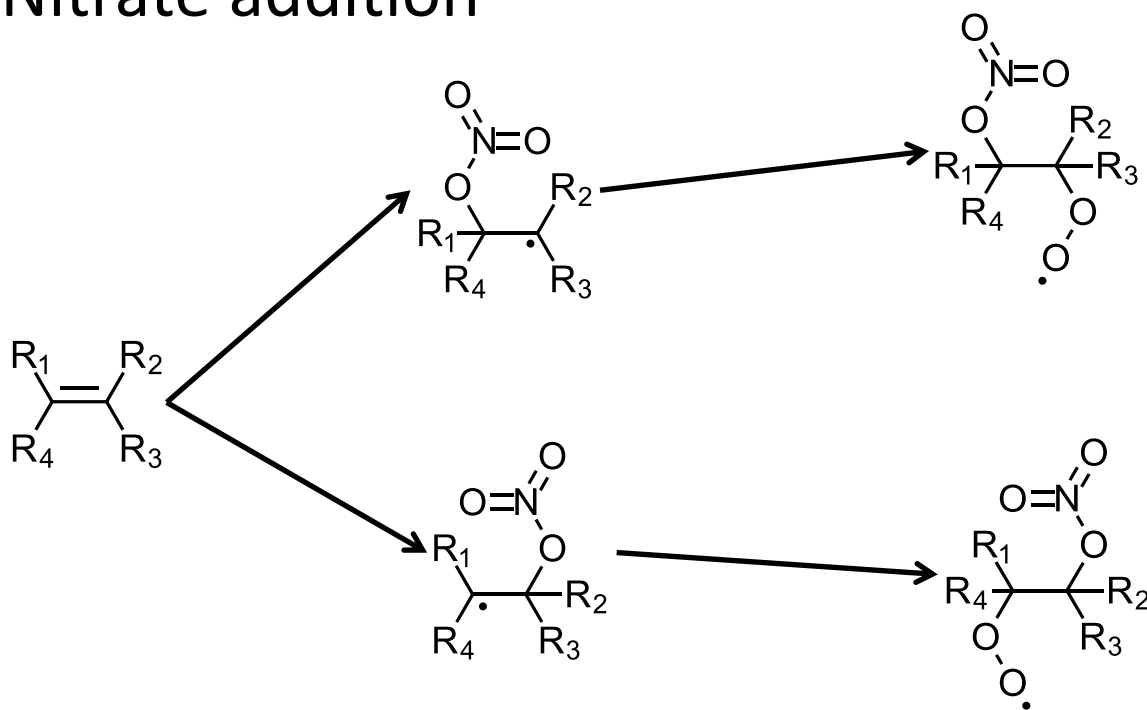
- Generally very low reactivity towards hydrogen abstraction
- Aldehydic C–H bonds are an exception to this rule



# Nitrate radical reactions



- Nitrate addition



- High reactivity towards activated alkenes
- Adds nitrate and epoxy functionality to alkenes

# Atomic chlorine



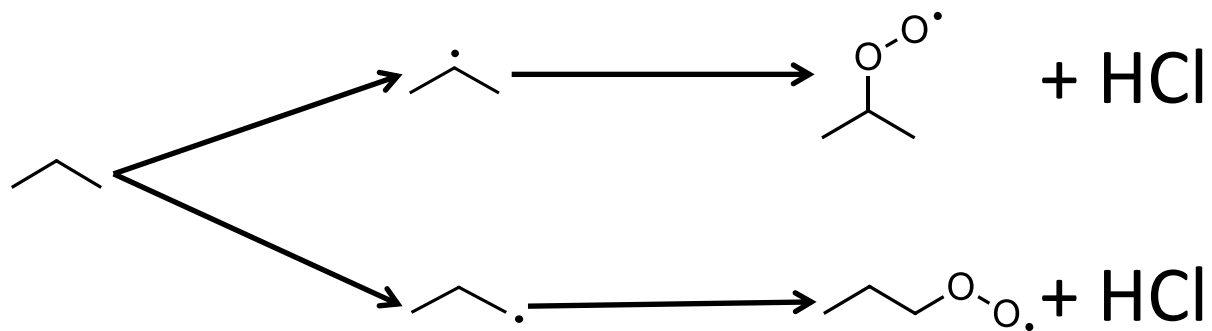
- Atomic chlorine is one of the most reactive oxidants known
- Typically, higher concentrations are seen in coastal environments
- This is not always the case with longer-lived precursors such as ClNO<sub>2</sub>
- Sinks: rapid reactions with organic compounds
- Ambient concentration:  $\sim 1 \times 10^4$  molecule cm<sup>-3</sup>



# Atomic chlorine reactions



- Hydrogen abstraction

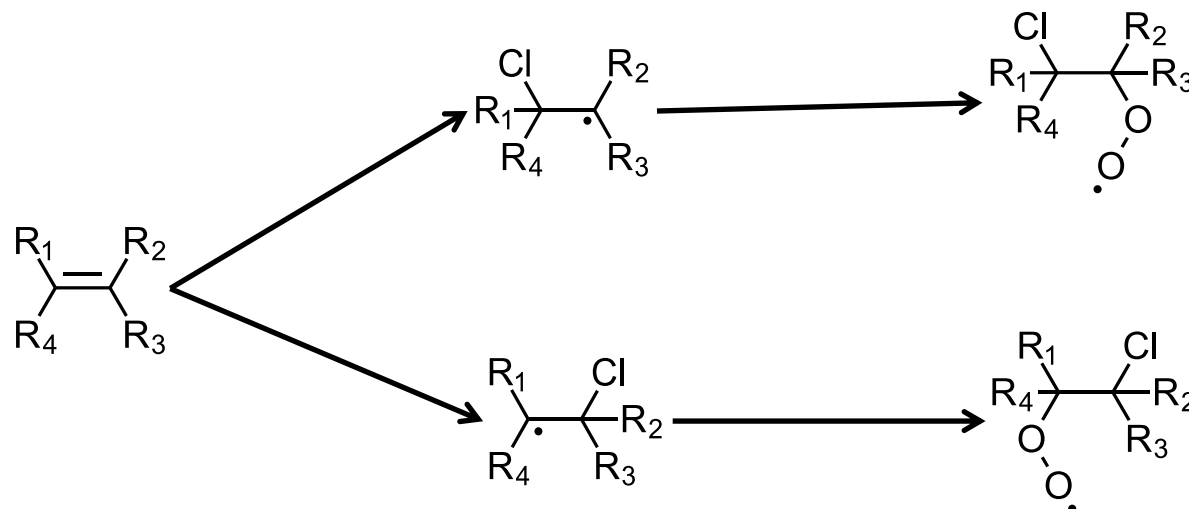


- Generally very rapid reactions with most types of C–H bonds

# Atomic chlorine reactions



- Chlorine addition



- Very rapid reactions with alkenes, less so with aromatics
- Adds chlorine functionality to alkenes structures

# Ozone



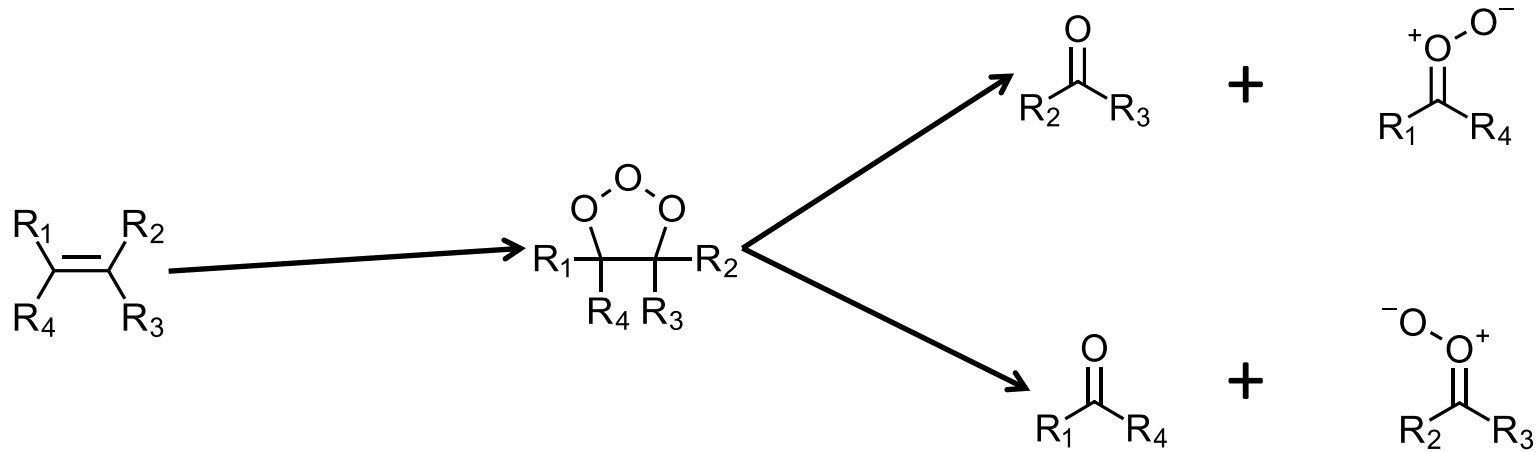
- Ozone is less reactive than other oxidants
- It is produced from the photochemical oxidation cycles in the atmosphere
- As a consequence,  $[O_3]$  can be much higher, especially in urban environments
- This means that reaction rates can be competitive with other oxidants
- Sinks: reactions with alkenes and NO
- Ambient concentration:  $\sim 7 \times 10^{11}$  molecule  $\text{cm}^{-3}$



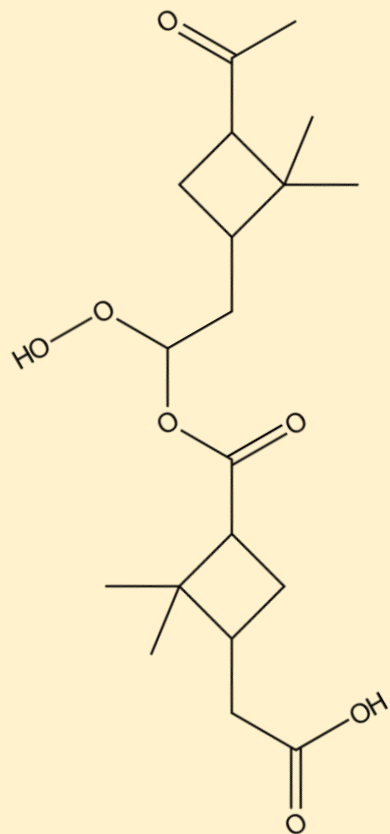
# Ozone reactions



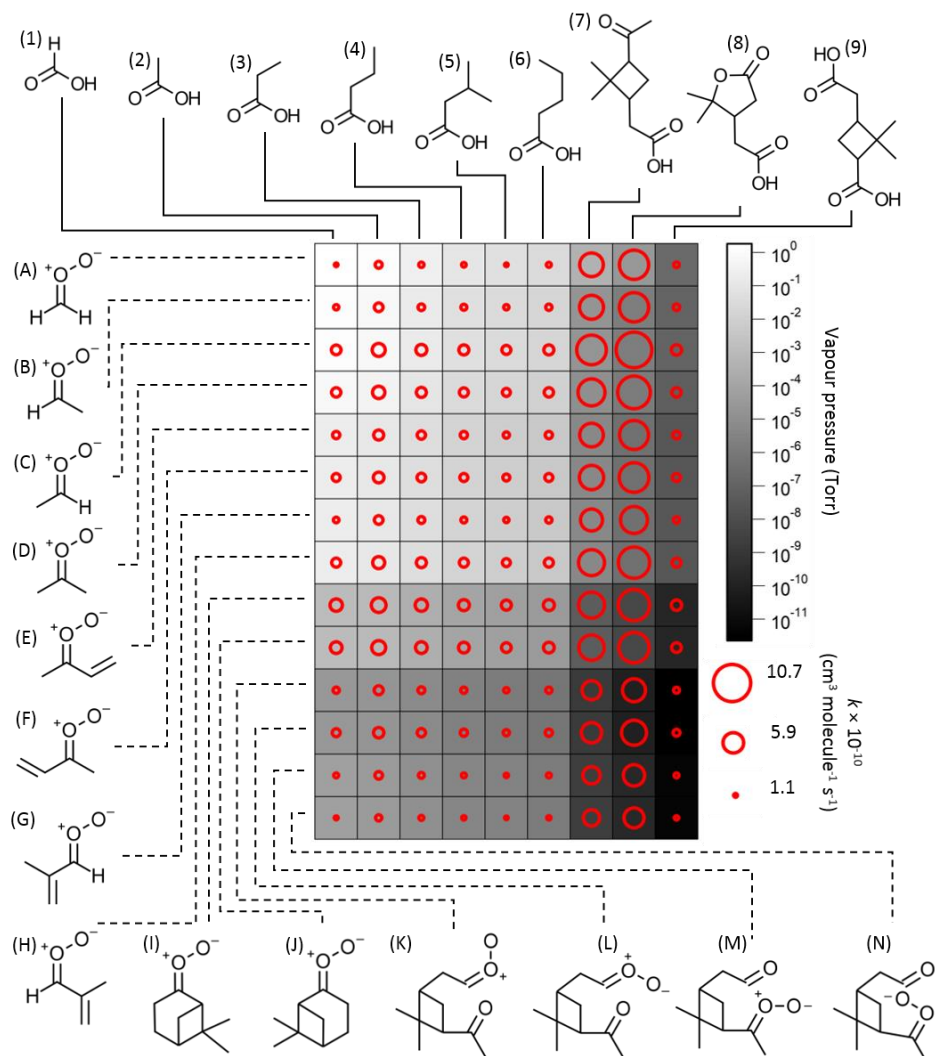
- Ozonolysis



# Criegee intermediates



**Above:** A low-vapour pressure hydroperoxide  
**Right:** A matrix of different Cl and acid combinations



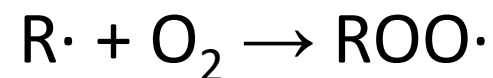
Source: Chhantyal-Pun et al. (2018)

- There are potentially lots of different Criegee intermediates and R(O)OH in the atmosphere
- These rate constants are really big
- It's a way of adding carbon and oxygen in one step
- **This could be a source of particles in the atmosphere**

# Peroxy radicals (RO<sub>2</sub>)



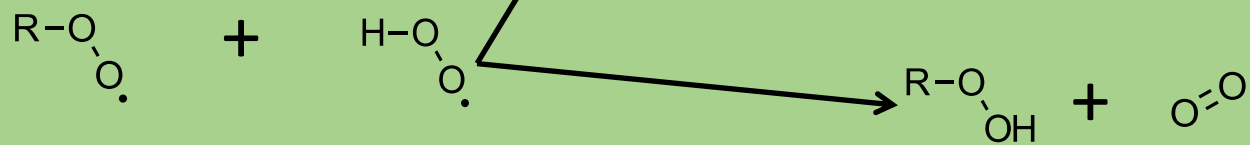
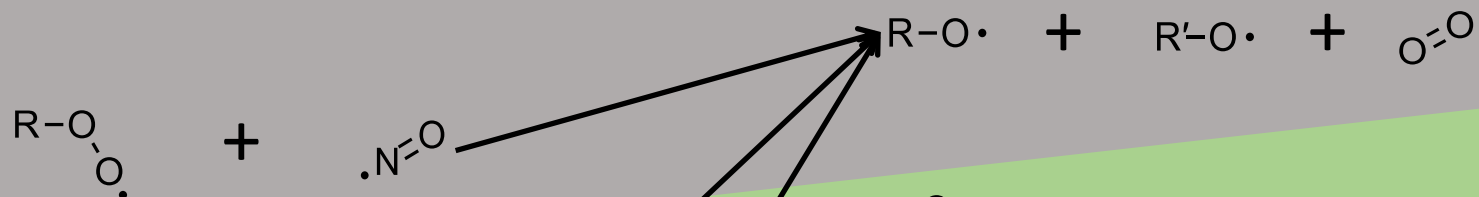
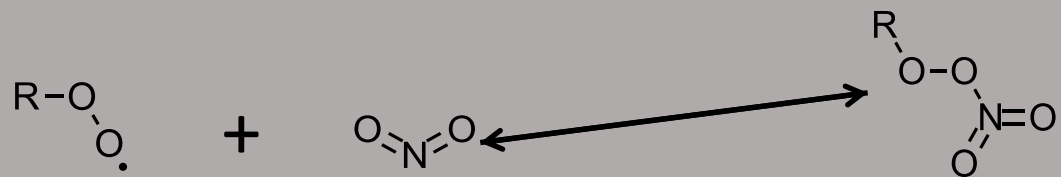
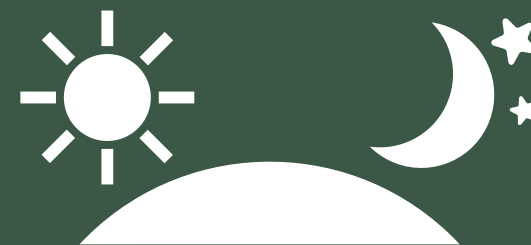
- Peroxy radicals are produced from the reaction of carbon-centred radicals with O<sub>2</sub>



- Carbon-centred radicals are produced during hydrogen abstraction and addition reaction mentioned earlier
- In urban environments, the sinks with NO<sub>x</sub> are very important
- In more remote environments reactions with other peroxy radicals become important
- For certain RO<sub>2</sub>, hydrogen shifts are very important



# Peroxy radical reactions



**URBAN**



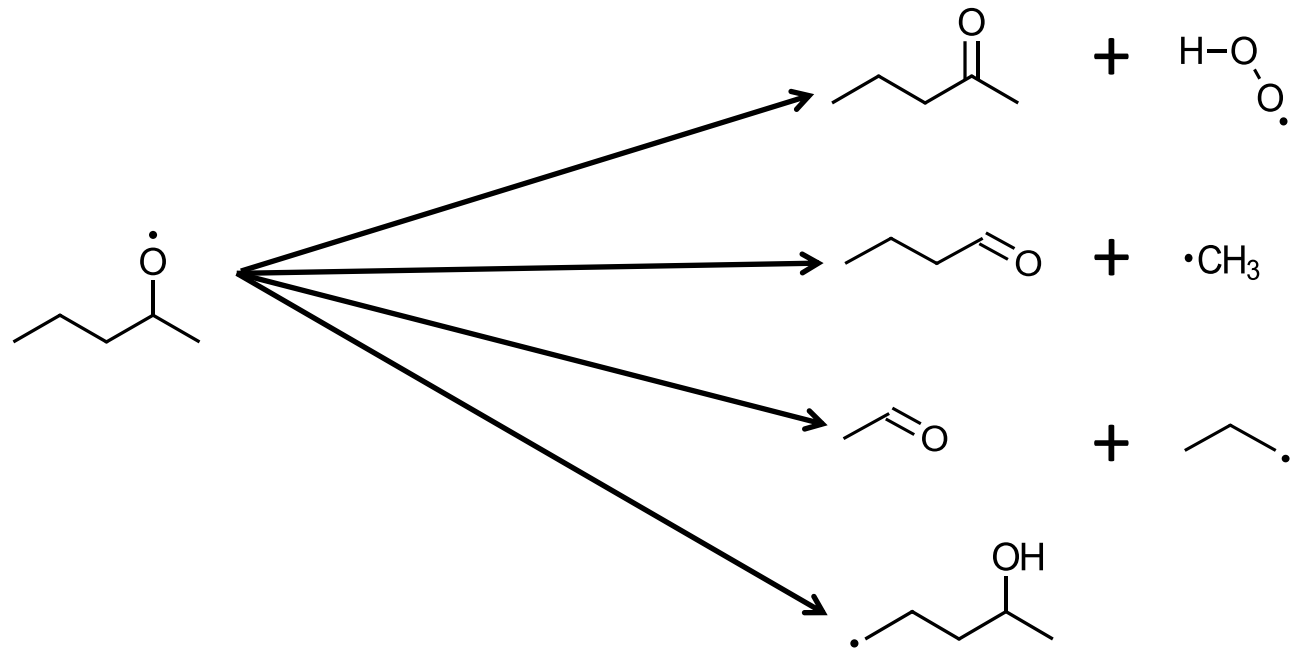
**REMOTE**

# Alkoxy radicals

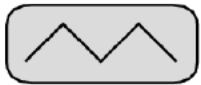


- Alkoxy radicals are short-lived intermediates that are produced from the reactions of  $\text{RO}_2$
- Because of their high reactivity towards  $\text{O}_2$ , their propensity for decomposition and rearrangement, their concentrations in the atmosphere are low
- Because of their ubiquity and the branching pathways available to alkoxy radicals, they are a major source of atmospheric complexity

# Alkoxy radical reactions

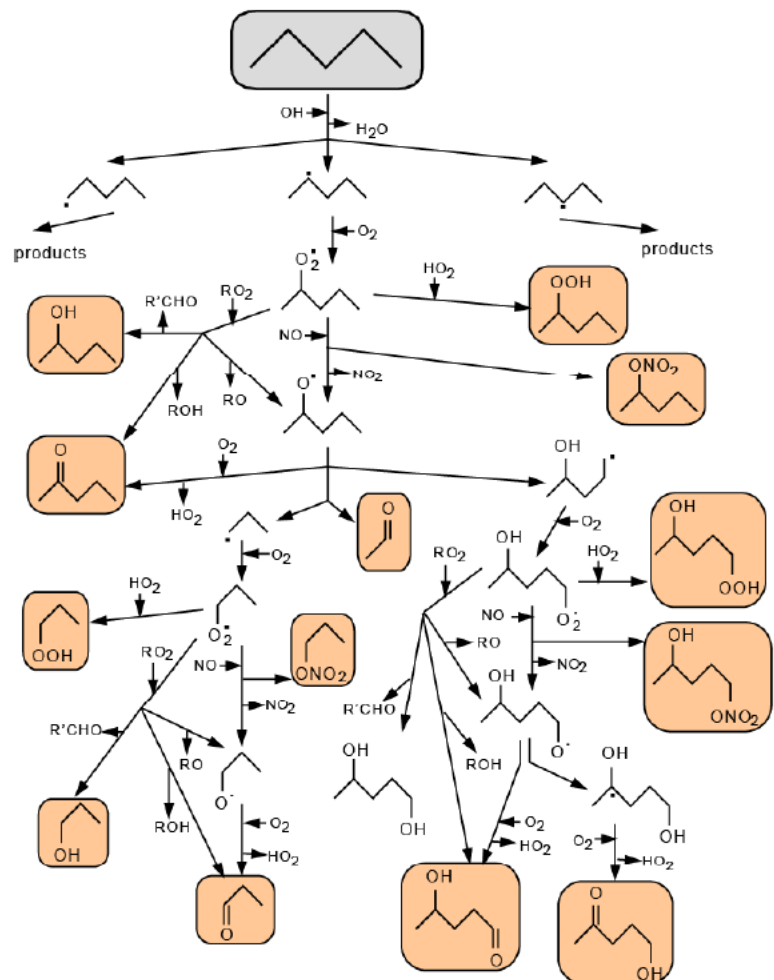






Parent  
Hydrocarbon

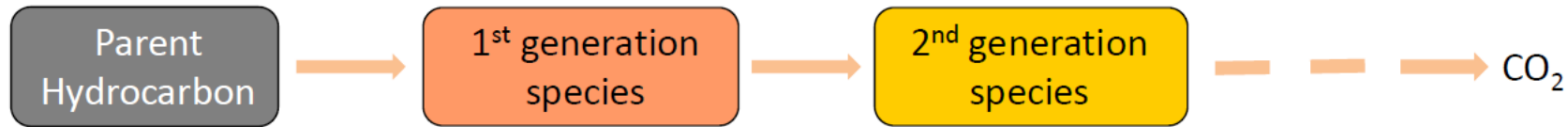
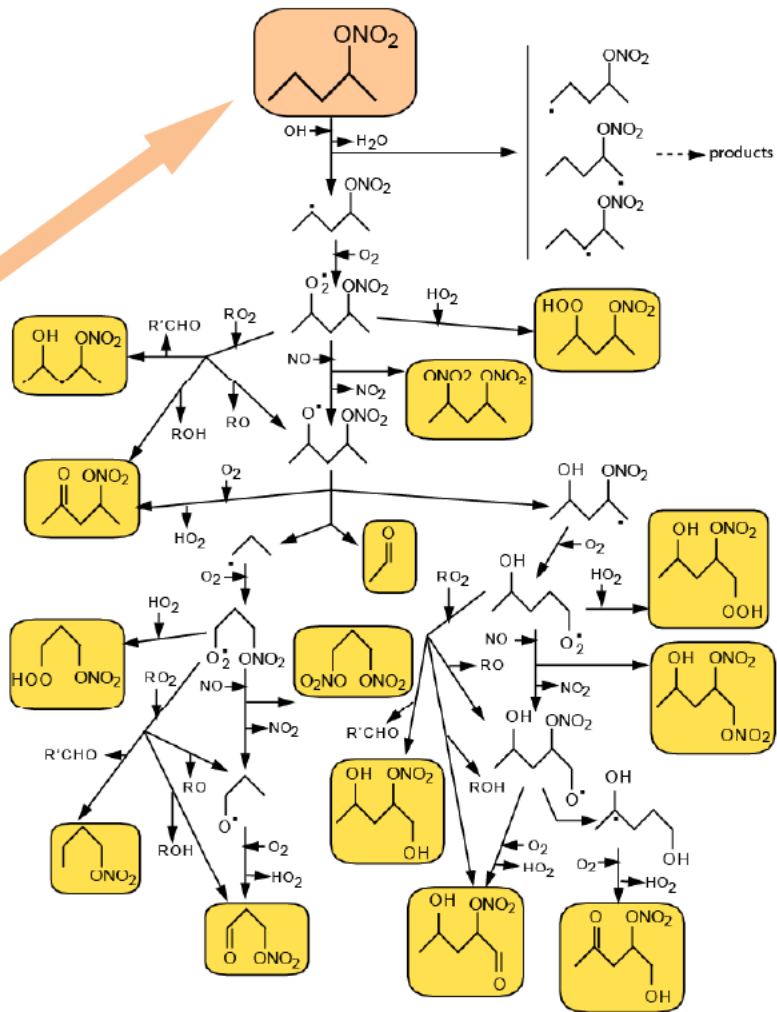
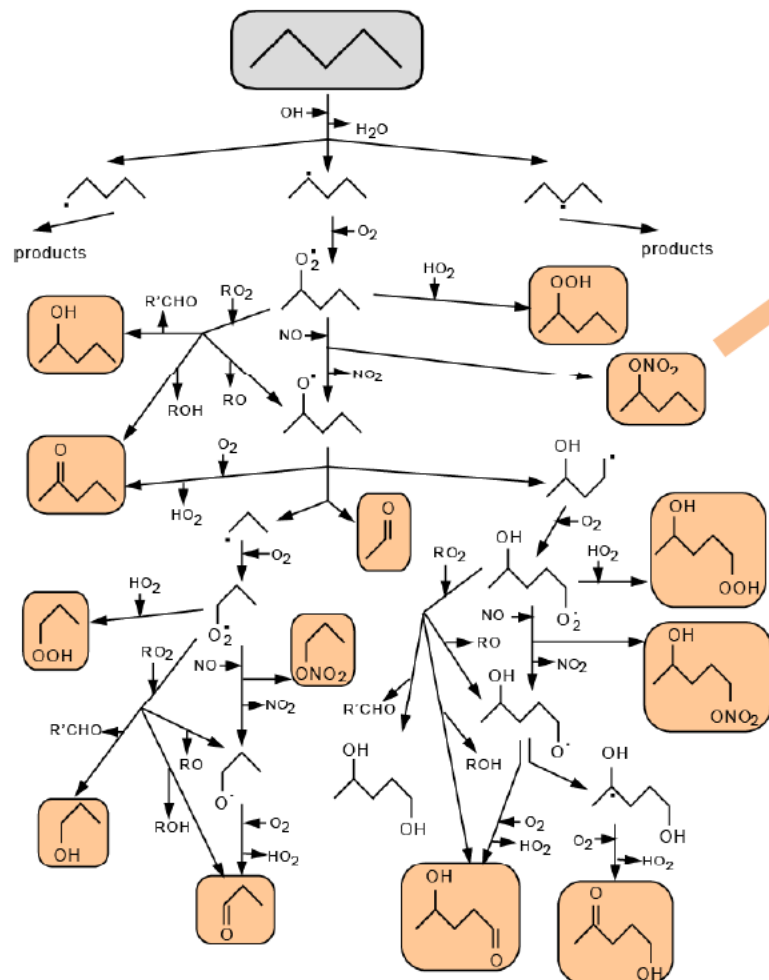




Parent  
Hydrocarbon

1<sup>st</sup> generation  
species

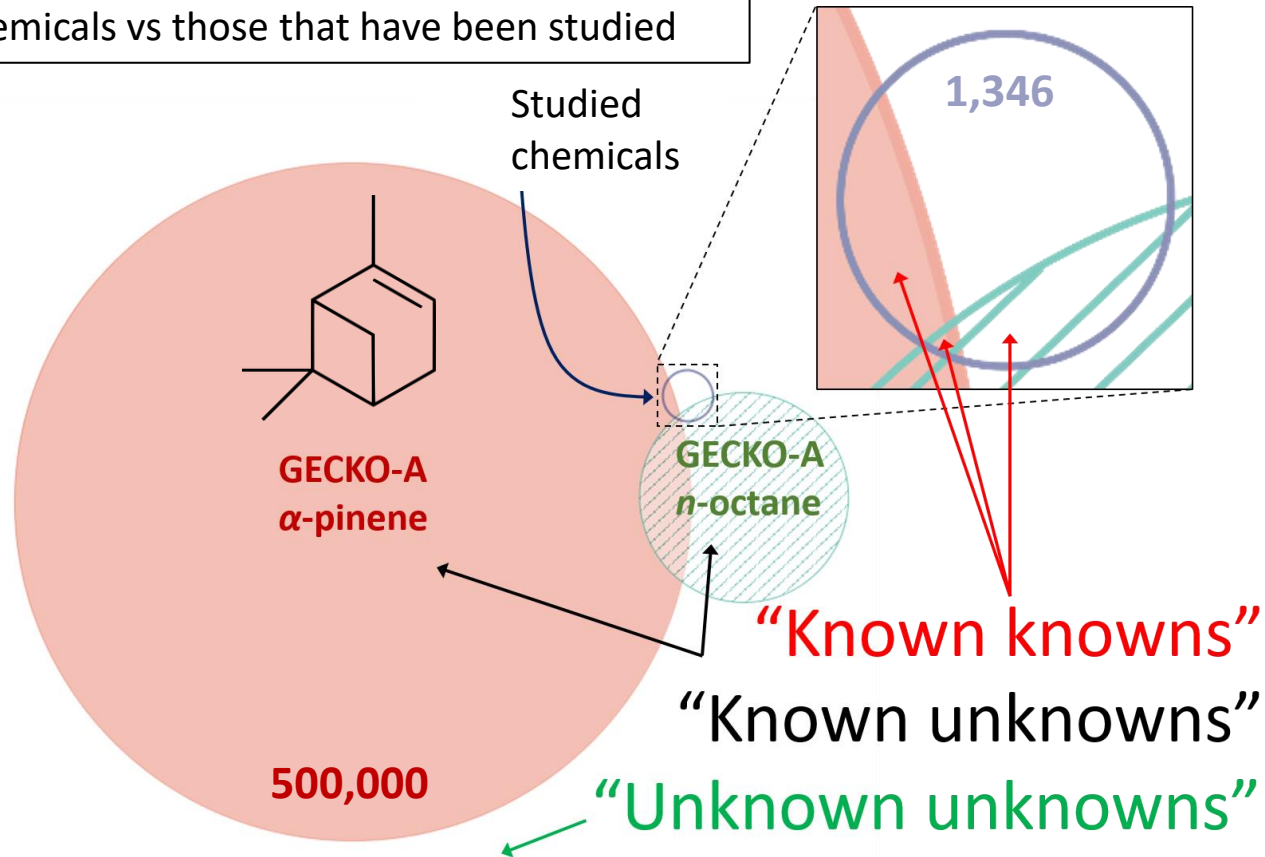
CO<sub>2</sub>





# The many unknowns of tropospheric chemistry

Area-proportional Venn diagram: number of products formed from two atmospheric chemicals vs those that have been studied



Source: McGillen et al. (2020)

- Far more chemicals expected to form in the atmosphere<sup>1</sup> than have been studied<sup>2</sup>
- Remember: these are just two primary emissions, what about the 1000s of others?<sup>3</sup>
- **It is surprising how much we don't know yet (reactivity, products, multiphase)**
- Why? Partly because of the huge numbers, partly because of the types of chemicals formed...

<sup>1</sup> Aumont, B. et al. Modelling the Evolution of Organic Carbon during Its Gas-Phase Tropospheric Oxidation: Development of an Explicit Model Based on a Self-Generating Approach. *Atmospheric Chem. Phys.* 2005, 5 (9), 2497–2517.

<sup>2</sup> McGillen, M. R. et al. Database for the Kinetics of the Gas-Phase Atmospheric Reactions of Organic Compounds. *Earth Syst. Sci. Data*, 2020, 12, 1203–1216.

<sup>3</sup> Goldstein, A. H.; Galbally, I. E. Known and Unexplored Organic Constituents in the Earth's Atmosphere. *Environ. Sci. Technol.* 2007, 41 (5), 1514–1521.

# Conclusions:

## Part 3 – Chemical reactions in the troposphere

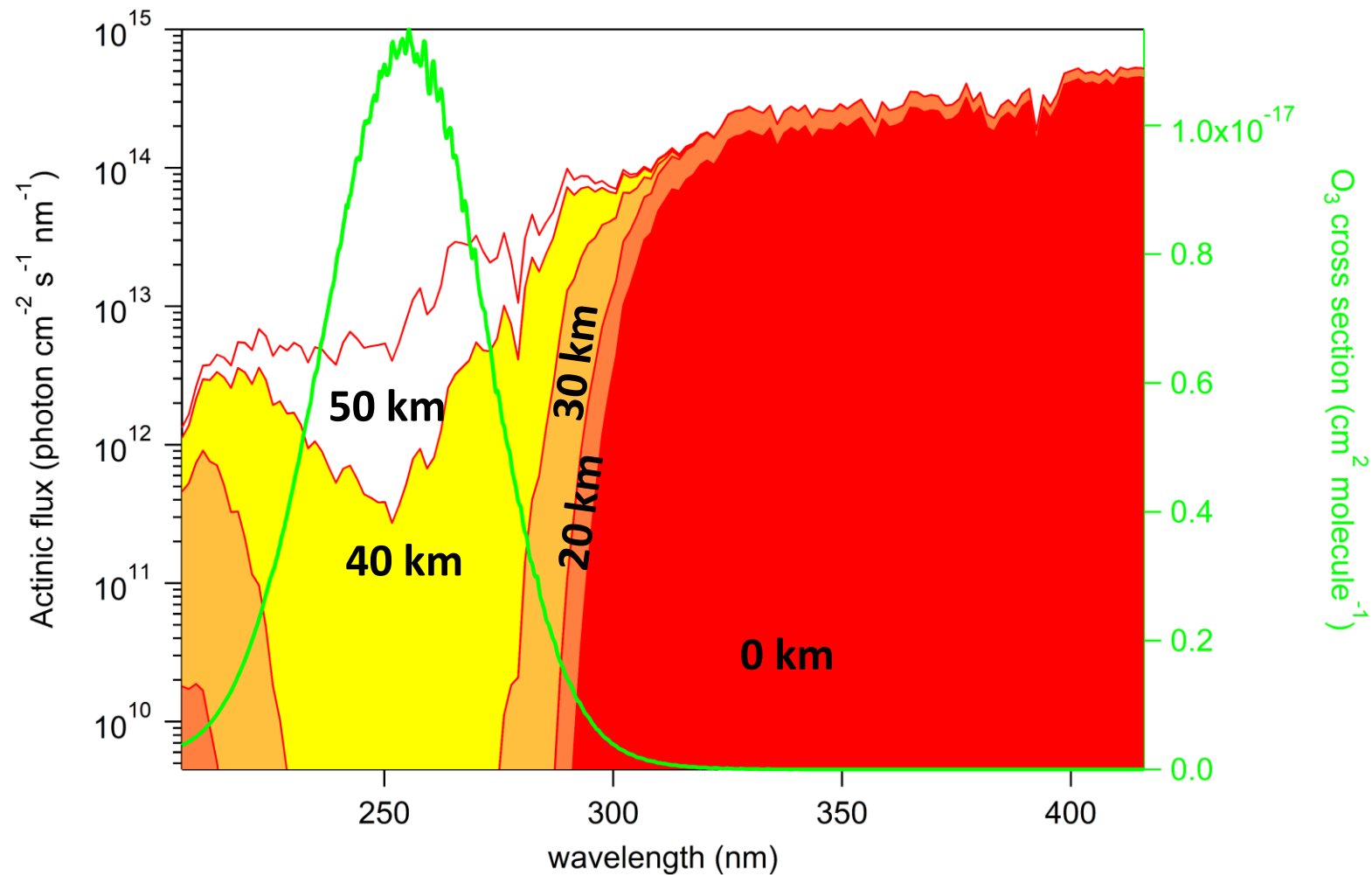
- The troposphere contains the largest variety of chemical reactions in the atmosphere
- It is an auto-catalytic cycle involving sunlight,  $\text{NO}_x$ ,  $\text{O}_3$  and VOCs
- These reactions clean the atmosphere, but they also produce harmful degradation products along the way
- The level of complexity that is generated from tropospheric chemistry is an example of a combinatorial explosion

# Part 4 – Chemical reactions in the stratosphere



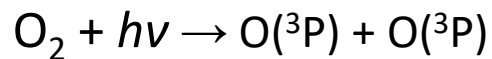
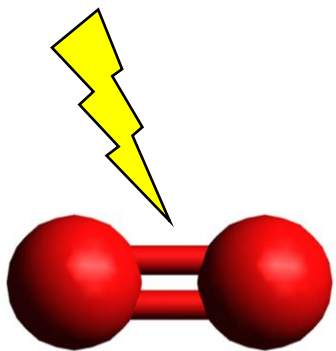
# Stratospheric ozone

- The ozone layer is one of the most important features of the atmosphere
- It is what attenuates most of the UV wavelengths in the troposphere
- Going above the ozone layer opens up the Herzberg continuum...





$$h\nu < 240\text{nm}$$



$$O \rightarrow 249 \times 2 \text{ kJ mol}^{-1}$$

From Hess's law:

$$\Delta H_{\text{rxn}} = \nu(\text{p})\Delta H_{\text{f}(\text{p})} - \nu(\text{r})\Delta H_{\text{f}(\text{r})}$$

$$= 498 \text{ kJ mol}^{-1}$$

$$= \underline{498,000 \text{ J mol}^{-1}}$$

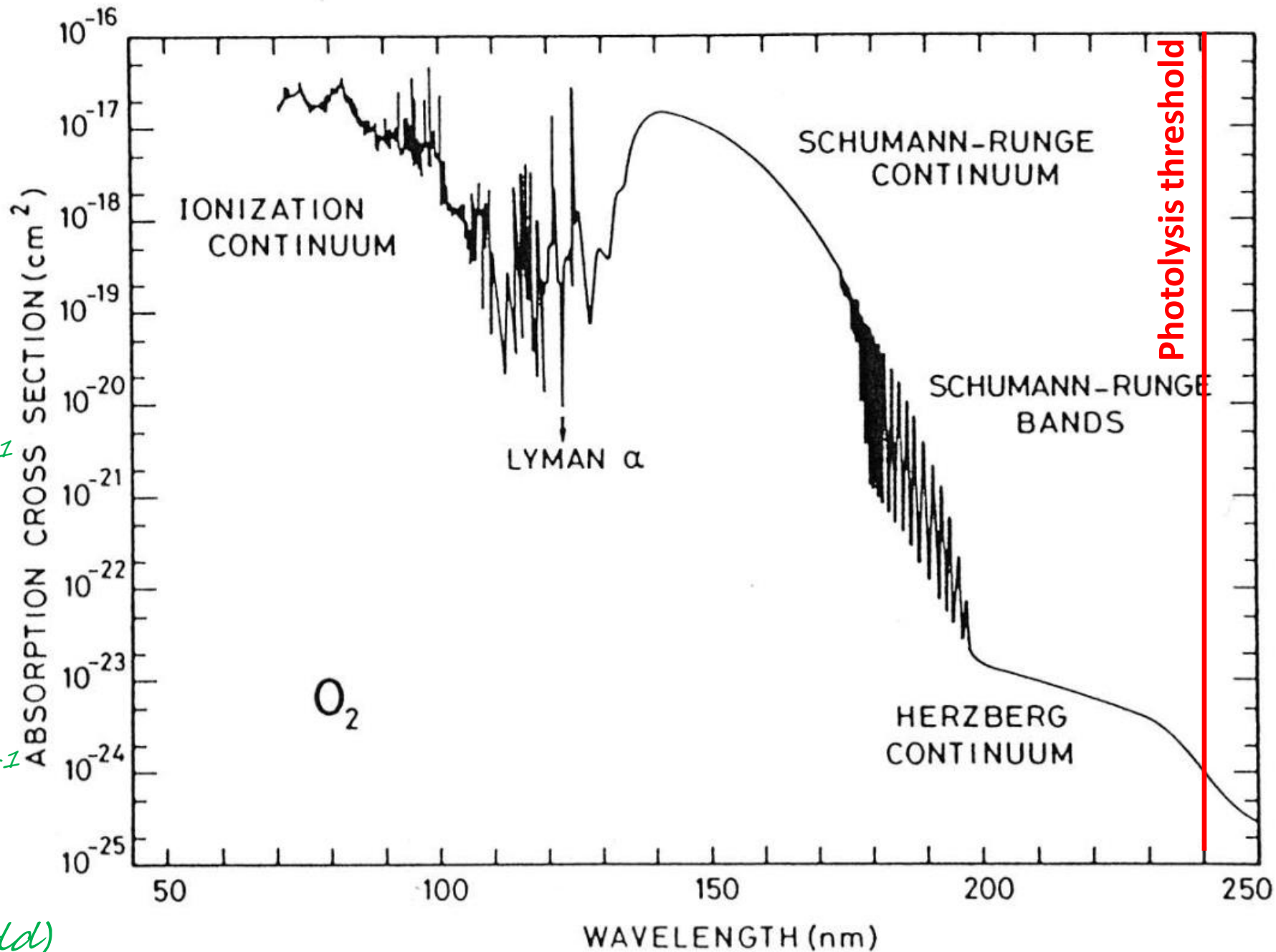
$N_A$

$$= \underline{8.27 \times 10^{-19} \text{ J photon}^{-1}}$$

$h \times c$

$$= \underline{240 \text{ nm}}$$

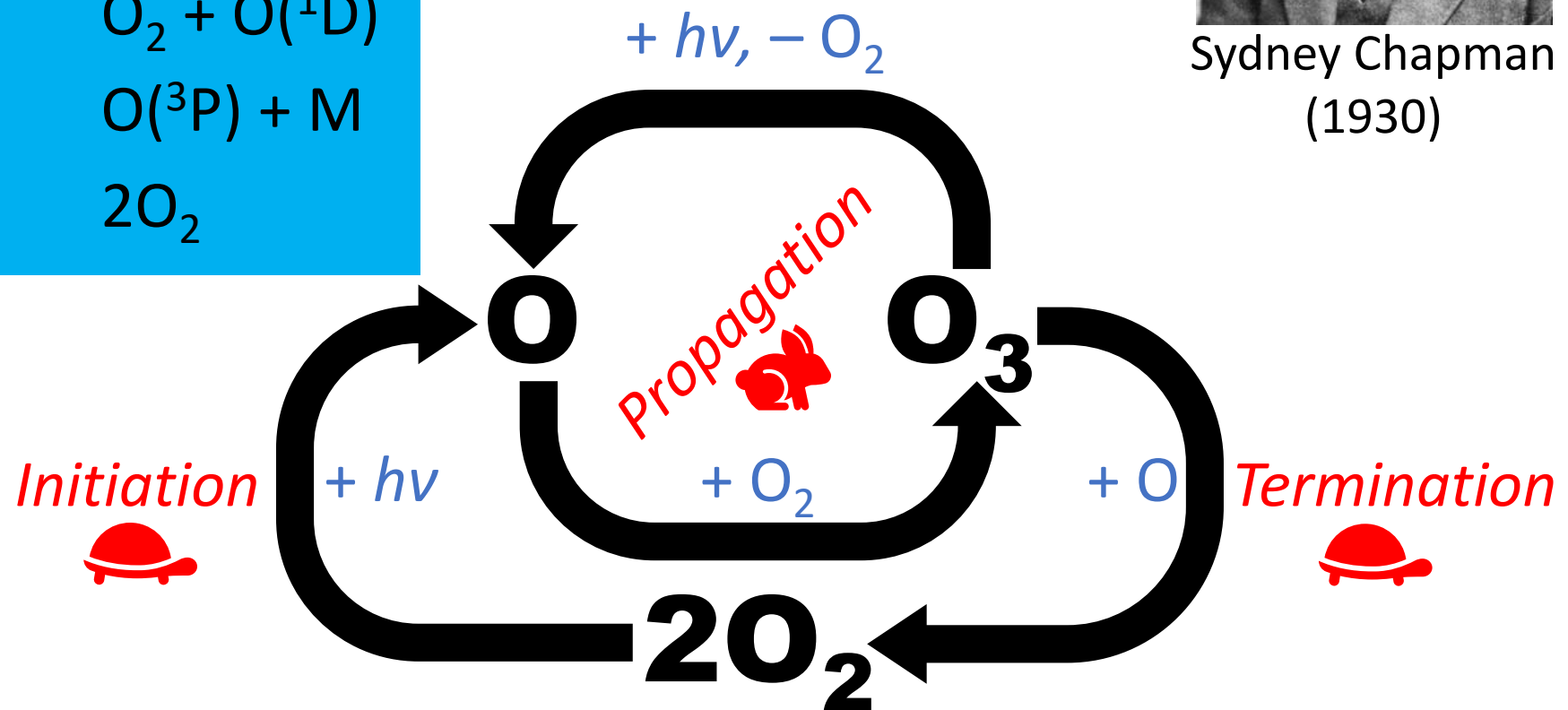
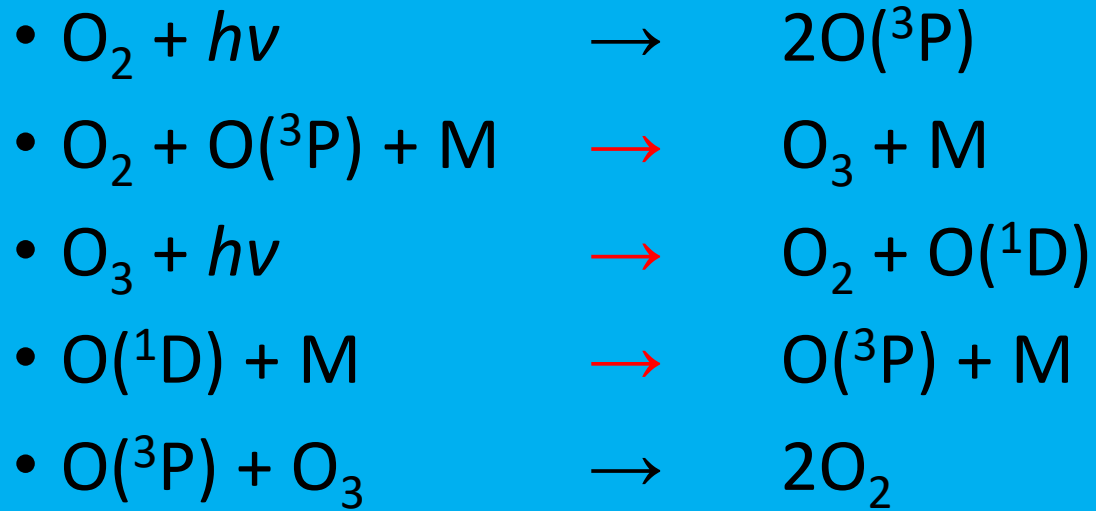
(the photolysis threshold)



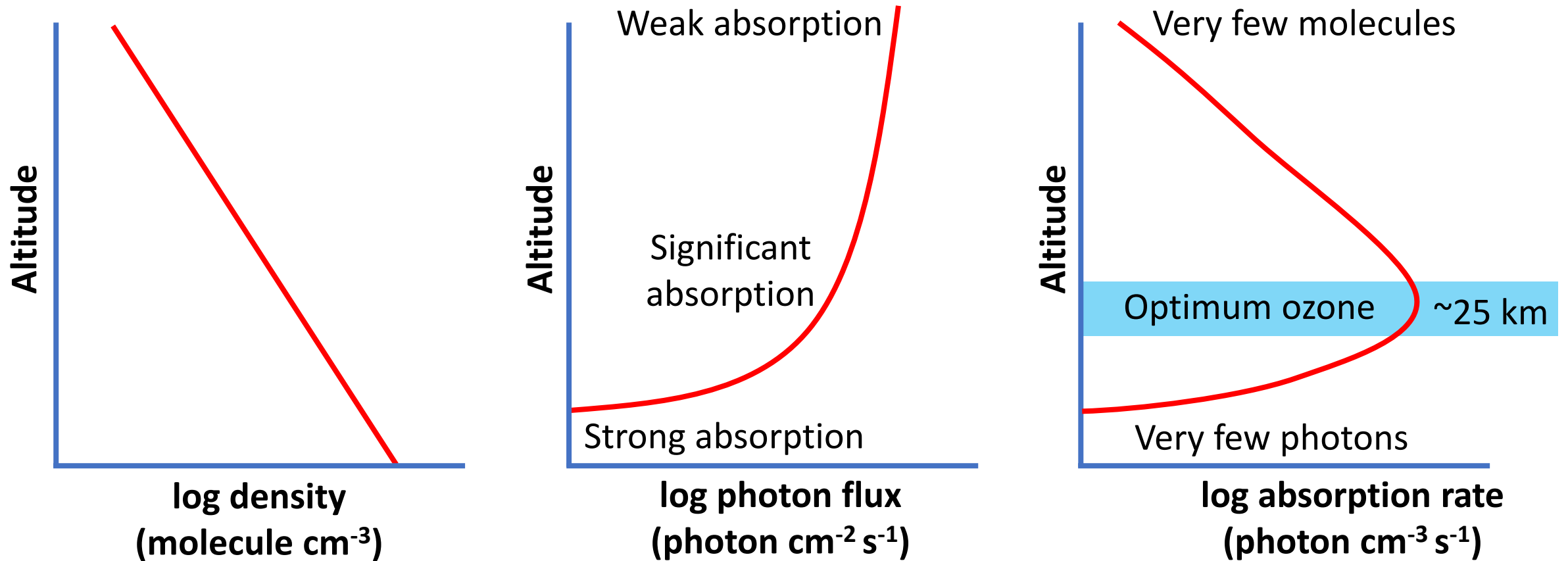
# The Chapman Cycle



Sydney Chapman  
(1930)



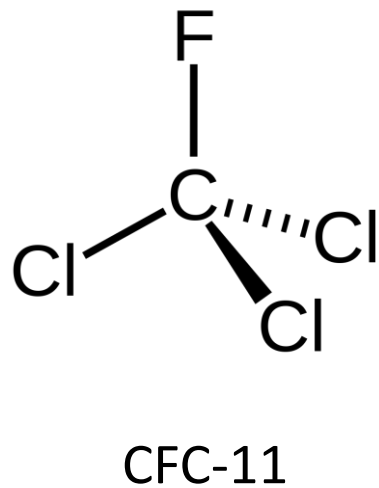
# The Chapman profile



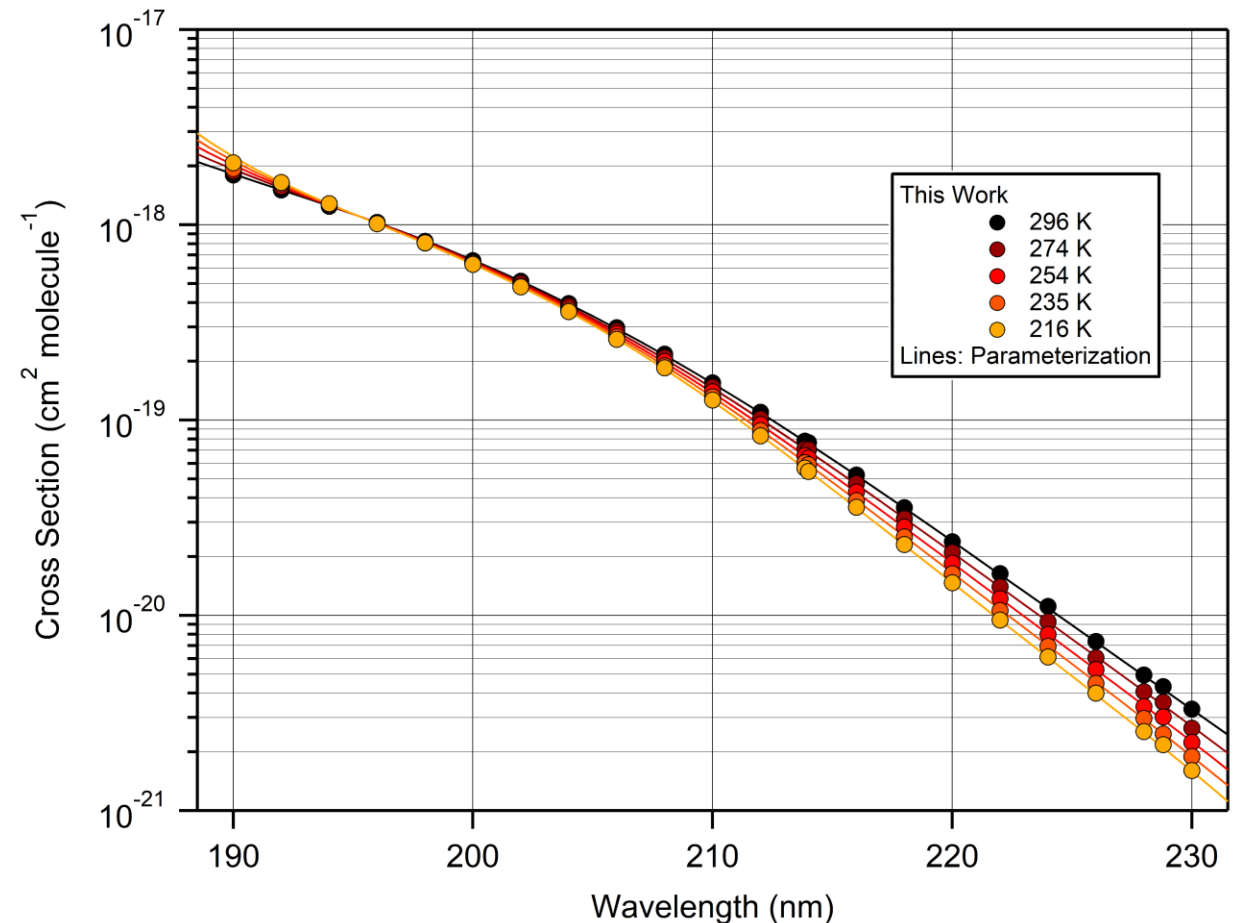
Ozone production is a **balance** between number density, photon flux and absorption cross sections of  $\text{O}_2$  and  $\text{O}_3$

# - Quiz -

- Remember all those reactions that are available to oxidize chemicals in the troposphere?
- What happens to this molecule?



- No reaction with any of the tropospheric oxidants
- No significant absorption in the actinic range





## The Nobel Prize in Chemistry 1995



James Lovelock  
(1972)



Paul J.  
Crutzen



Mario J.  
Molina



F. Sherwood  
Rowland

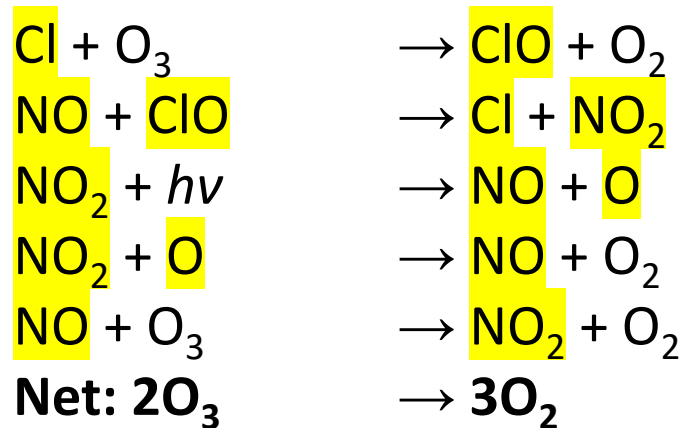
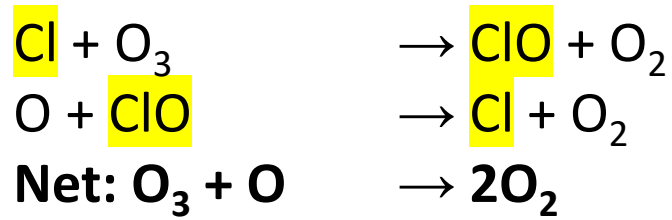
- CFC-11 ( $\text{CFCl}_3$ ) was observed in the air samples of James Lovelock in 1972
- Molina and Rowland determined that these molecules survive for decades/ centuries
- This indicates upward transport
- Together with Crutzen they anticipated the threat that this presents to the stratosphere...

# Catalytic ozone destruction

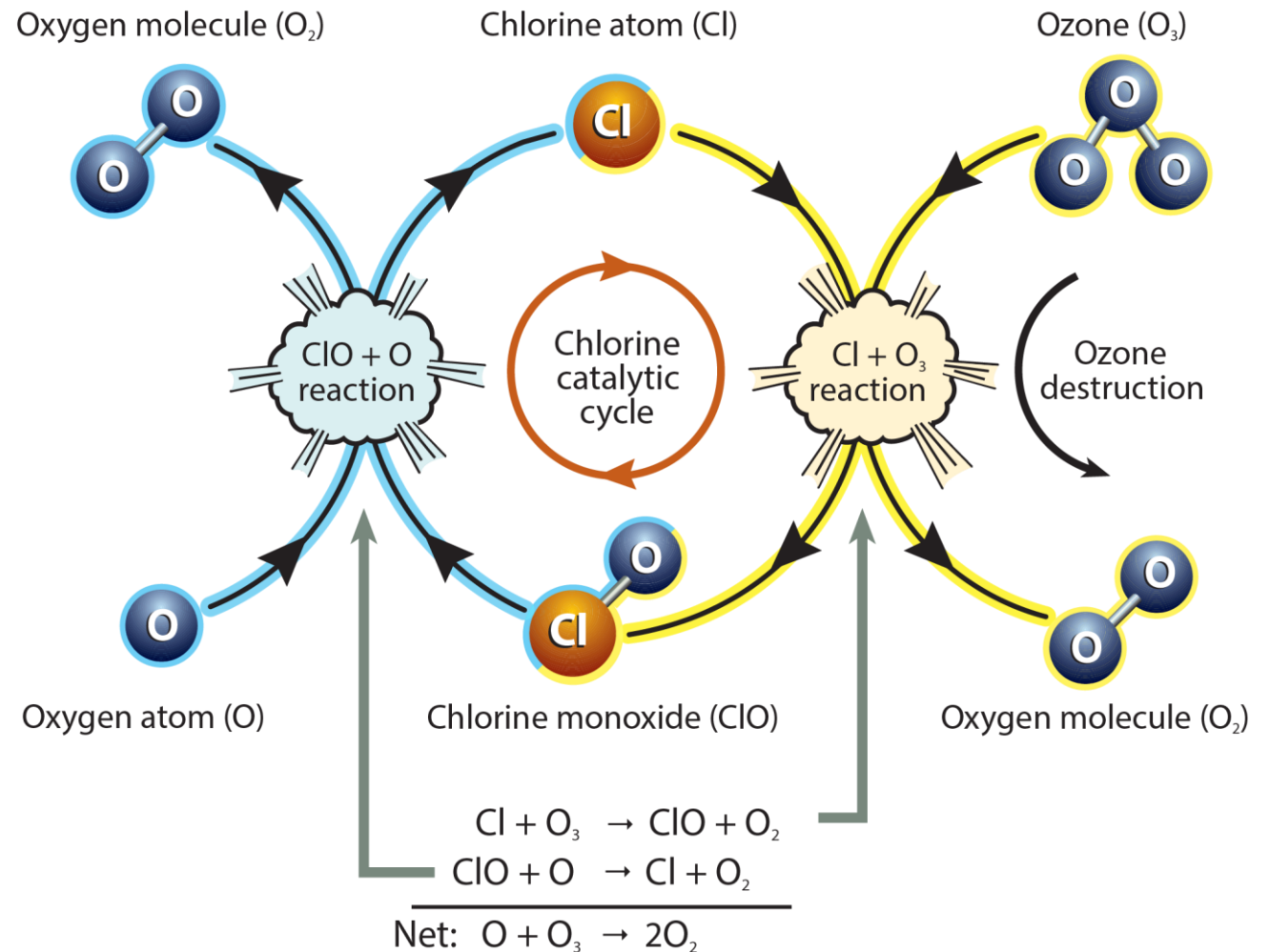
## Initiation:



## Propagation:

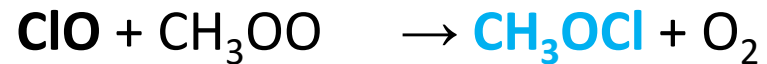
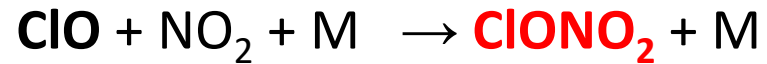


## Ozone Destruction Cycle 1



# Catalytic ozone destruction

## Termination:

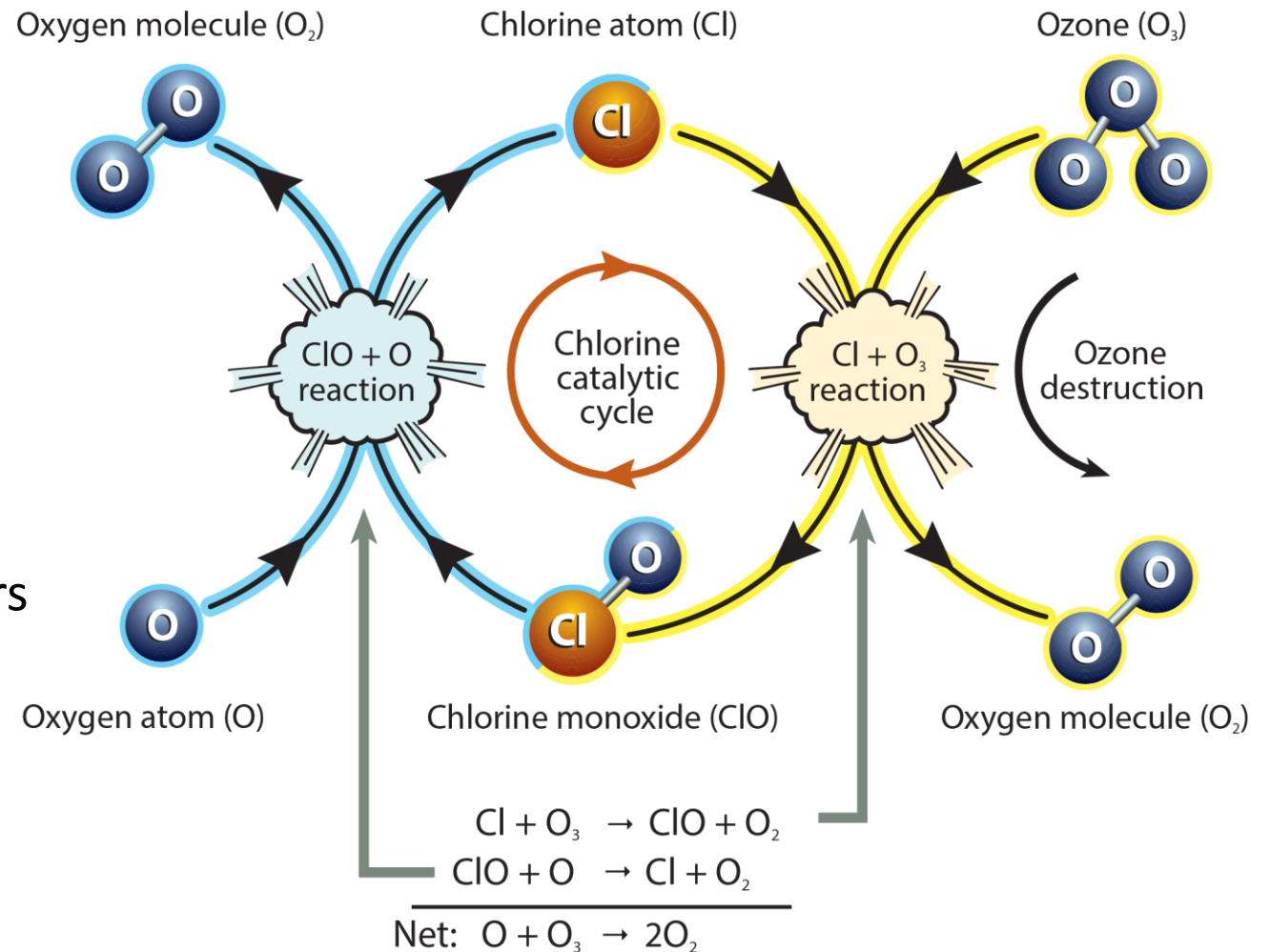


**Cl and ClO:** active chlorine

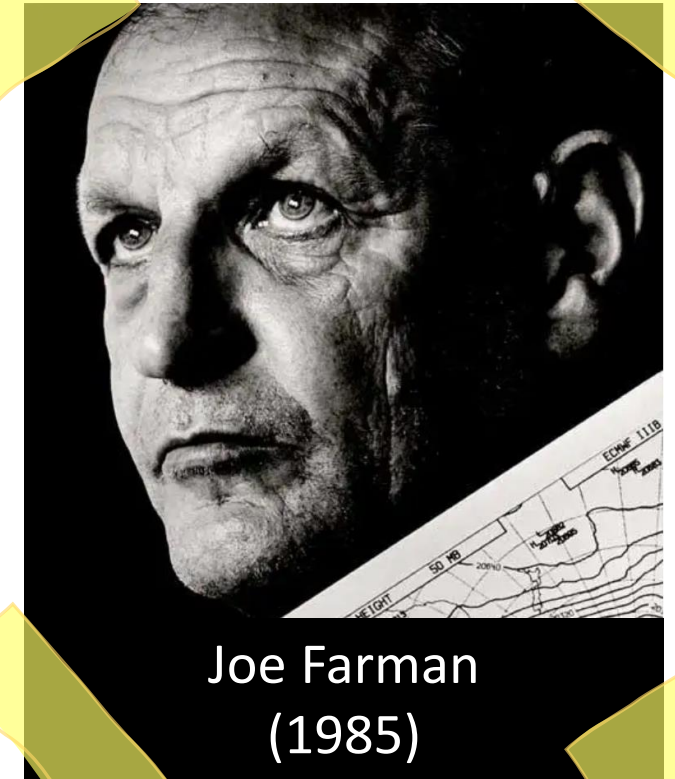
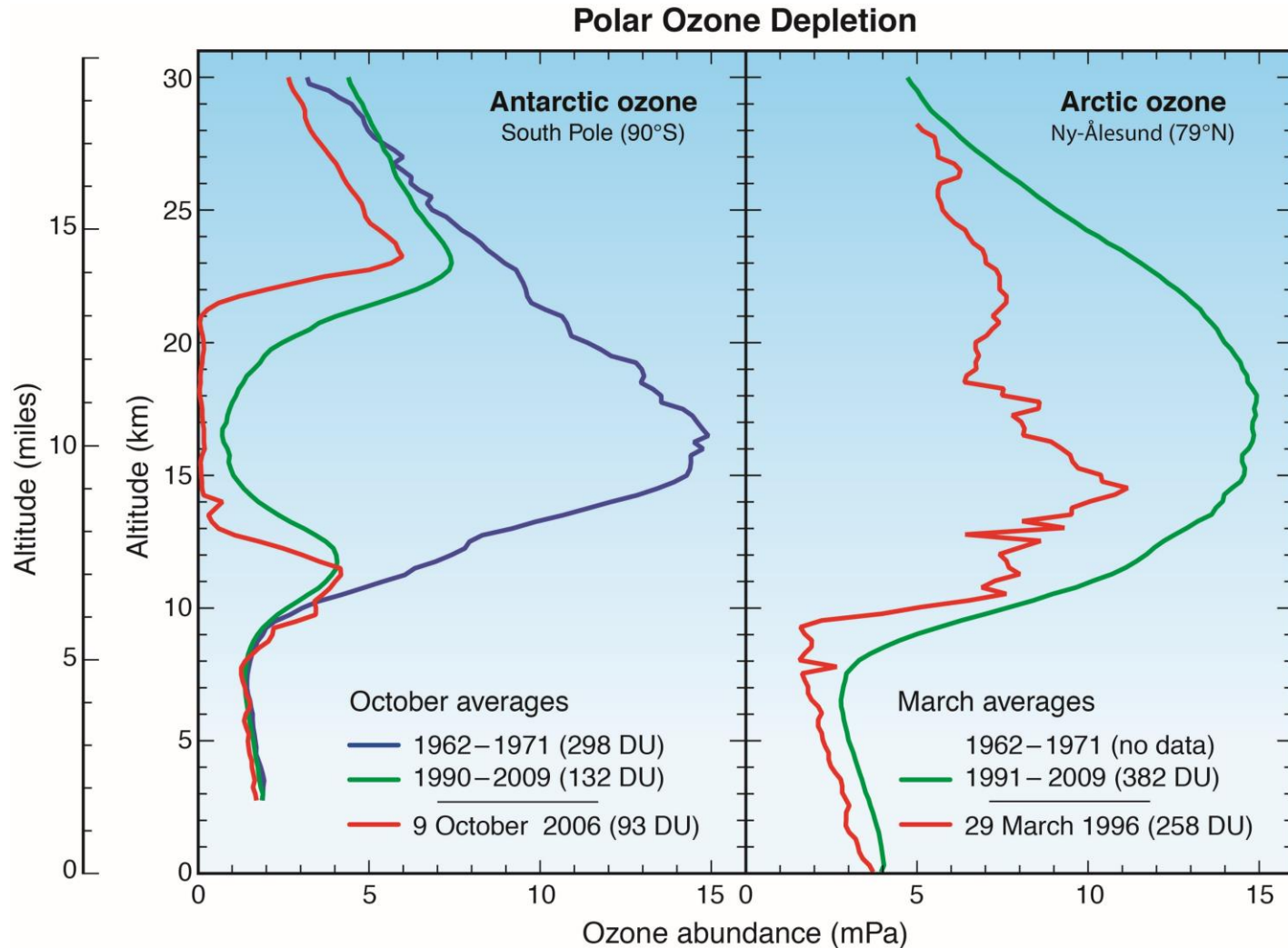
**HCl and ClONO<sub>2</sub>:** important Cl reservoirs

**HOCl and CH<sub>3</sub>OCl:** less important Cl reservoirs

## Ozone Destruction Cycle 1



# The ozone hole



Joe Farman  
(1985)

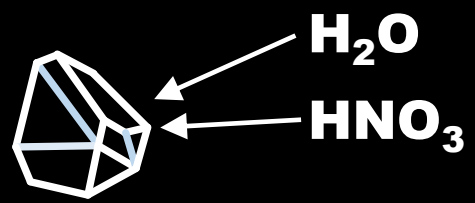
- Observation of the ozone hole was very sudden and unexpected
- This led to a rapid international effort to understand this process further



# In the polar night

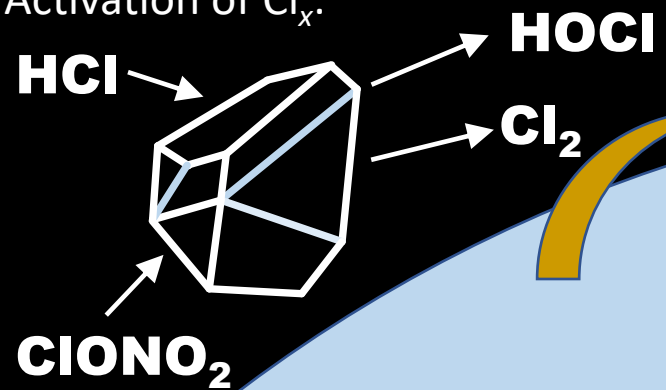
1

Formation of polar stratospheric clouds:



2

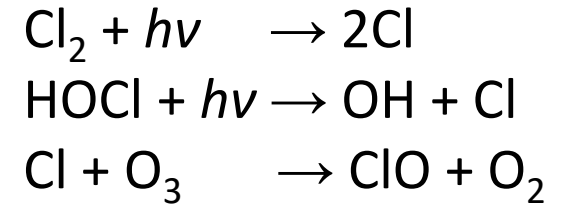
Activation of  $\text{Cl}_x$ :



# In the polar spring

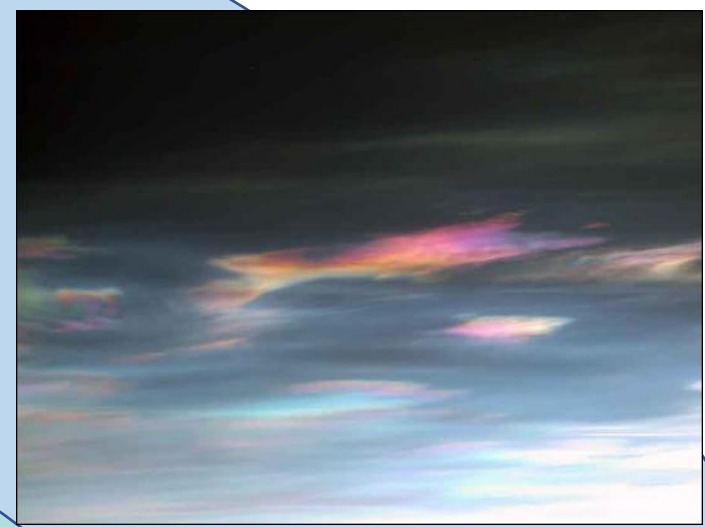
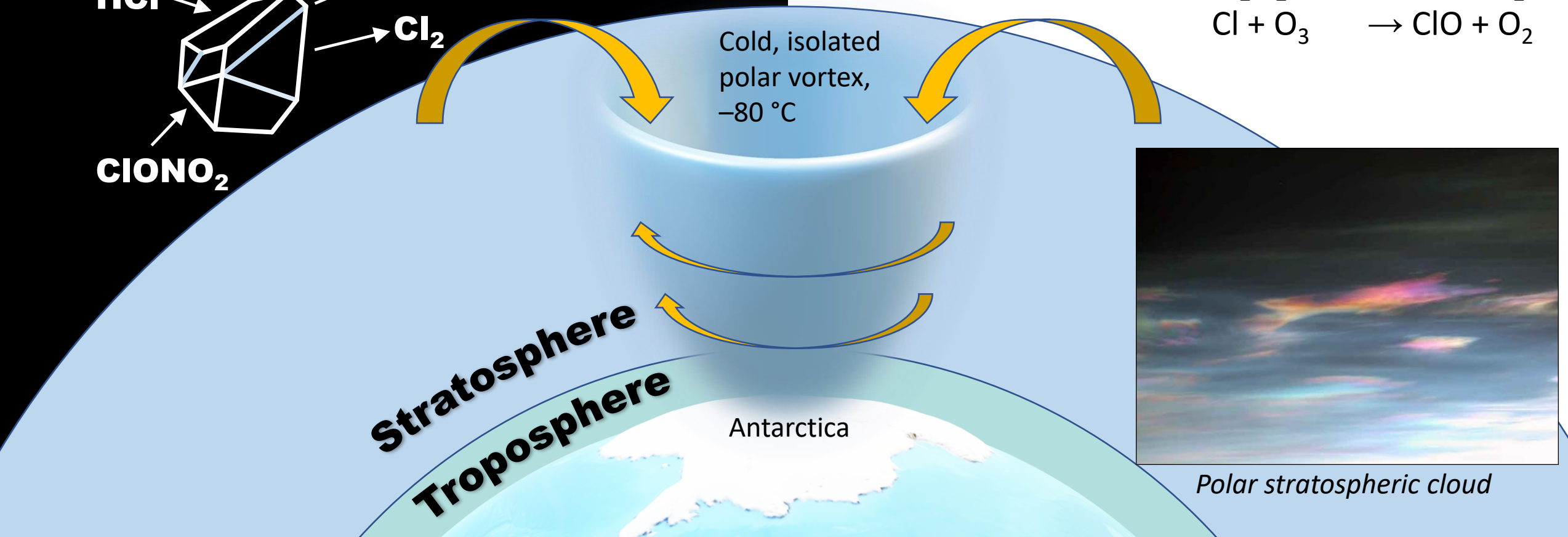
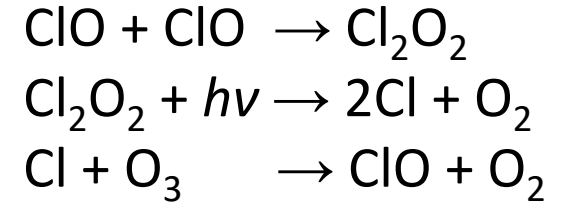
3

Formation of  $\text{ClO}_x$ :



4

Catalytic  $\text{O}_3$  depletion:



Polar stratospheric cloud

# Conclusions:

## Part 3 – Chemical reactions in the stratosphere

- As you get higher into the atmosphere, shorter UV wavelengths become accessible
- This allows for the photolytic production of O<sub>3</sub>, and the destruction of very stable compounds such as CFCs
- This provides a source of chlorine into the stratosphere which has a catalytic effect on ozone depletion
- Society's rapid response to ozone depletion is often cited as a success story in atmospheric chemistry
- But there are other ways of looking at this...
- How will society respond to other future environmental threats?

# Approaches to studying the atmosphere

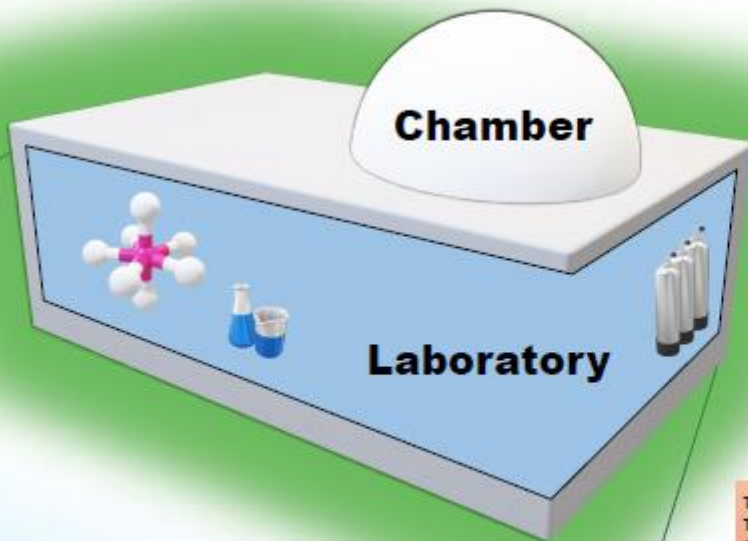
- Studying the atmosphere is a big job
- This requires many different activities
- This can be viewed metaphorically as a 3-legged stool
- In the field, we can observe what the atmosphere is composed of
- In the lab, we can understand the individual processes to create and destroy species
- In models, we can couple our fundamental understanding from the lab with our field observations





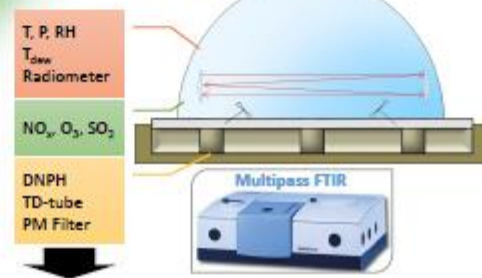


A view of the HELIOS platform at sunrise



# HELIOS platform

HELIOS instrumentation



Scenes inside the HELIOS laboratory

