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Article

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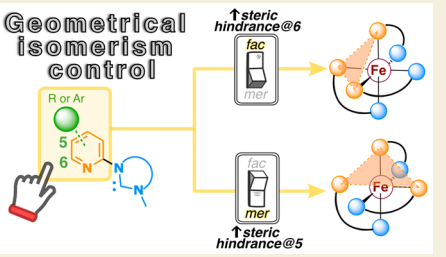
ABSTRACT: The stereoselective synthesis of geometrical iron(II) complexes bearing azine-NHC ligands is described. Facial and meridional selectivity is achieved as a function of the steric demand of the azine unit, with no remarkable influence of the carbene nature. More specifically, meridional complexes are obtained upon selecting bulky 5-mesityl-substituted pyridyl coordinating units. Unexpectedly, increase of the steric hindrance in the α position with respect to the N coordinating atom results in an exclusive facial configuration, which is in stark contrast to the meridional selectivity induced by other reported α-substituted bidentate ligands. Investigation of the structure and the optical and electrochemical properties of the here-described complexes has revealed the non-negligible effect of the *fac*/*mer* ligand configuration around the metal center.



KEYWORDS: *Bidentate ligands, computational calculations, fac/mer isomers, iron complexes, NHC ligands*

INTRODUCTION

The stereoselective synthesis of octahedral complexes with unsymmetrical bidentate ligands $\mathbf{M(A\text{^}B)_3}$ keeps attracting considerable interest due to the intrinsic relationship between molecular structure and resulting properties. Indeed, it is well-known that coordination of this type of ligands often results in a mixture of facial and meridional stereoisomers in a statistical 1:3 ratio (Figure 1a).[1] However, these isomers can have different photophysical, electrochemical, and magnetic properties,[2−5] with appealing applications in OLEDS,[2,6,7] catalysis[7,8] or supramolecular chemistry.[9−11] Considering that isolation of these isomers can be extremely challenging, especially in the case of labile complexes, attaining control over the *fac*/*mer* stereoisomerism is thus of great importance.

In the most general case, the synthesis of facial isomers is enthalpically-driven, which is at the origin of the so-called trans effect,[12] while that of the meridional congeners is entropically-driven.[13] Nevertheless, systematic deviations from these statements are commonly encountered due to steric reasons or the presence of additional intra-/intermolecular interactions.[1] Interestingly, strategies based on solvation effects,[10,14] supramolecular interactions,[15−18] the use of preorganized platforms,[11,19−23] and exposure to external stimuli such as heat[24−26] or light[24,27] have been shown to influence the *fac*/*mer* ratio.

On the other hand, the development of photoactive iron complexes is currently a hot topic within the scientific community. Despite the challenging photophysics of typical polyimine Fe complexes, charge transfer-excited states with lifetimes ranging from hundred of picoseconds[28,29] up to few nanoseconds[3−32] have been achieved. These impressive results are the consequence of judicious ligand designs, with N-heterocyclic carbenes (NHC),[29,30] cyclometallating[31] units, or amides[32] playing a pivotal role in the creation of sufficiently high ligand fields.

During our recent research works in this field, we have reported a series of bidentate pyridyl-NHC iron (II) complexes with lifetimes comparable to those of their tridentate analogues while having one NHC unit less.[5] Although the metal−ligand interaction was improved, the asymmetry of the ligands led to the obtention of geometrical isomers (Figure 1b). In the case of **C0** comprising pyridine-imidazol-2-ylidene ligands, a 1:14 *fac*/*mer* ratio was obtained.[33] In stark contrast to typical Fe^{II} complexes, no lability was observed for **C0**. However, isolation of these geometric isomers was not attained. As a result, molecular modeling was selected to provide some insights into the photophysical properties of these distinct species, especially regarding their excited state properties. Interestingly, these results showed that, upon electronic excitation, the facial configuration leads to slower relaxation kinetics, resulting in longer excited state lifetimes.[33,34] Subsequent preparation of a pure facial complex was possible by means of the tripodal $\mathbf{C_3}$ ligand **L1**.[23] Nevertheless, the influence of the connecting scaffold and the increased rigidity on the final properties of the facial isomer cannot be ruled out. Increase of the π-conjugation

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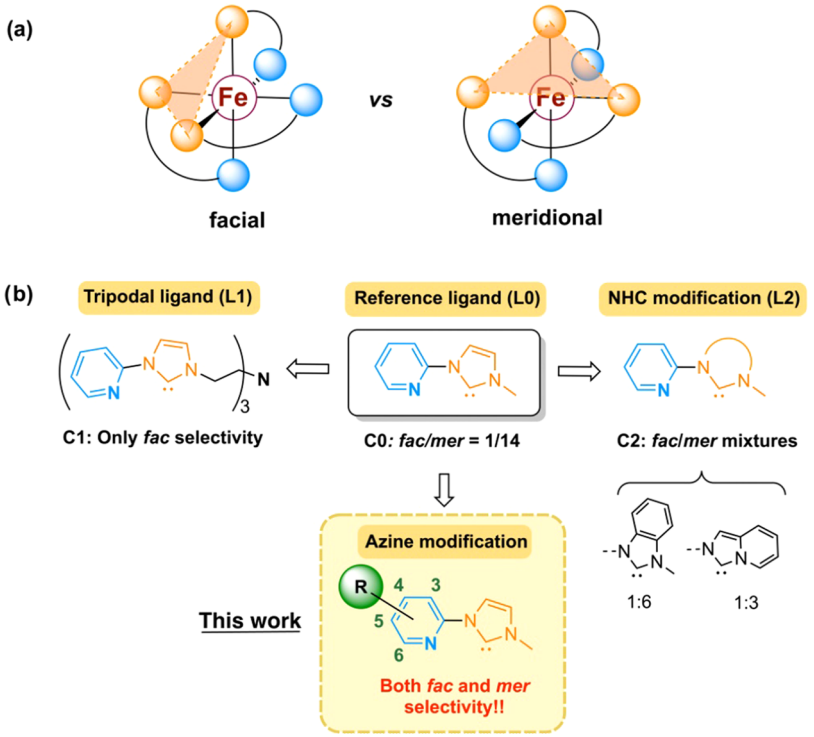
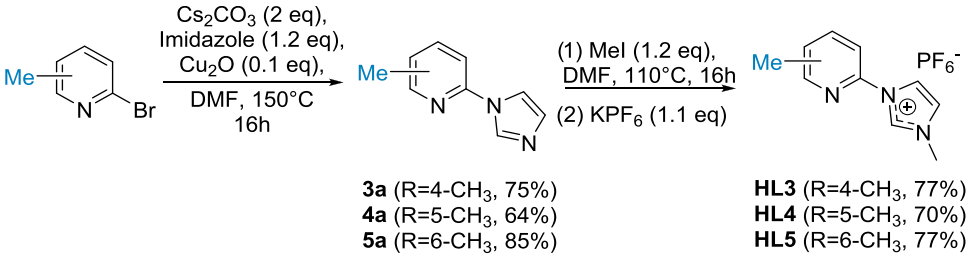
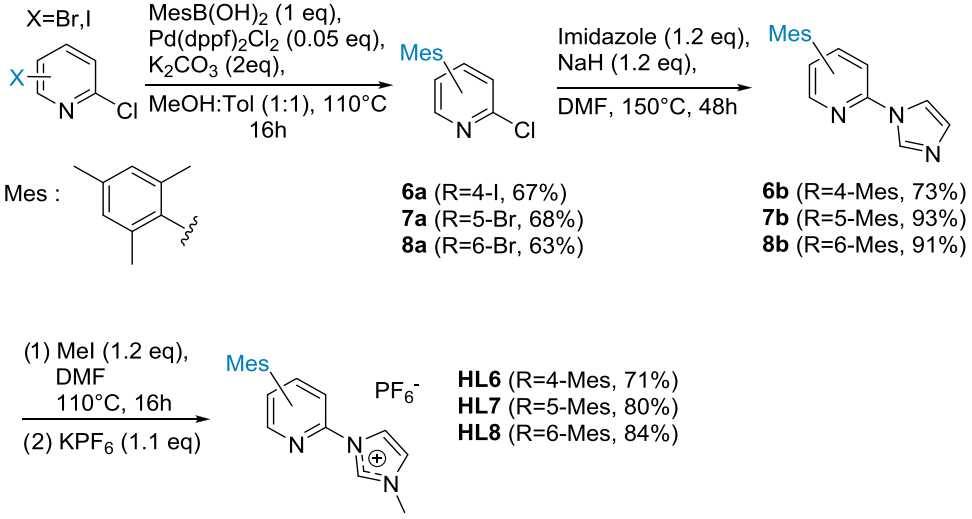


Figure 1. (a) Schematic representation of facial and meridional configurations for iron(II) complexes bearing asymmetric bidentate ligands. (b) Overview of the explored ligand designs and their impact on the *fac*/*mer* ratio of the corresponding iron(II) complexes.

Scheme 1. Synthesis of Ligand Precursors HL3−HL5



Scheme 2. Synthesis of Ligand Precursors HL6−HL8



at the NHC unit (**L2** ligands) afforded opposite modulations of the opto-electronic properties depending on the localization of the benzannulated ring.[35] Noteworthy, the increased steric demand of the ligands leads to a modification of the stereoselectivity toward higher *fac*/*mer* ratios with respect to parent **L0**, highlighting the importance of a rational ligand

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design. Prompted by these results and the need to exert precise control on the coordination sphere, we explore in this contribution the effects of modifying both the substitution nature and pattern at the azine unit on the final complex configuration, which turns out to be pivotal for *fac*/*mer* isomer selectivity (Figure 1b). In particular, an unexpected facial selectivity is provided by 6-substituted pyridyl-/quinolyl-/quinoxalyl-NHC ligands which is, to the best of our knowledge, unique for this type of sterically demanding bidentate ligands.

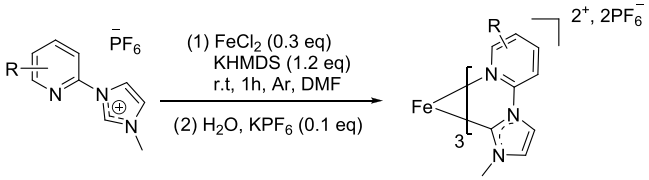
RESULTS AND DISCUSSION

Two different ligand precursor series were initially prepared for the introduction of two levels of steric hindrance. A methyl (Me) (**HL3**−**HL5**) was selected as the low level to minimize the inductive or resonance effects (Scheme 1). A mesityl (Mes) group (**HL6**−**HL8**) was chosen instead as the high level not only due to its high steric demand but also because of the additional possibility of promoting π-stacking interactions (Scheme 2).[36] Substitution was performed at positions 4, 5, and 6 at the pyridine moiety for both series. Position 3 was not functionalized since a weaker interaction with the metal ion could be anticipated due to the concomitant loss of coplanarity upon such ligand modification.[4]

In all cases, pyridyl-imidazoliums were selected for the preparation of py-NHC bidentate ligands. Me-based precursors **HL3**−**HL5** were obtained in good yields by an Ullmann coupling of imidazole on the corresponding methylated 2-bromopyridine, followed by a quaternization with iodomethane in DMF at 110 °C and final metathesis with potassium hexafluorophosphate (Scheme 1). As for Mes-based precursors **HL6**−**HL8**, their synthesis was conceived from halogenated 2-chloropyridines. After the initial selective introduction of the mesityl group by means of a Suzuki coupling, the imidazole unit was introduced via a S_N^{Ar} mechanism at the Cl-substituted pyridine position. Upon azole quaternization with MeI and I^- to PF_6^- counterion metathesis, the target precursors were obtained in good yields (Scheme 2).

Based on our previous works with unsubstituted ligands, the complexation of **HL3**−**HL8** precursors was performed using $FeCl_2$ as iron source, KHMDS as the base, and DMF as solvent (Table 1). With the exception of **L8**, successful ligand coordination yielded **C3**−**C7** in *ca.* 45%, regardless of the ligand nature. Nevertheless, remarkable differences were obtained concerning the *fac*/*mer* selectivity. Within the Me-based series, no significant influence was observed for **L3** and **L4**, with the methyl group at 4 and 5 positions, respectively, obtaining **C3** and **C4** with a near statistic 1:3 *fac*/*mer* ratio (entries 1 and 2). However, the presence of the methyl group adjacent to the pyridinic N atom in **L5** nicely resulted in the exclusive formation of the facial isomer *fac*-**C5** despite its non-negligible steric hindrance toward the metal ion coordination (entry 3). This result was completely unexpected since previously reported 6-methyl-2,2′-bipyridine complexes resulted in meridional complexes.[37] Moreover, the complex remains in its low-spin configuration, which also differs from other similarly substituted facial complexes.[38,39]

Table 1. Effect of the Ligand Steric Demand on Fe^{II} Complex Formation



entry	ligand precursor	yield (%)[a]	*fac*/*mer* ratio[b]	complex
1	**HL3** (4-Me)	44	25:75	**C3**
2	**HL4** (5-Me)	44	20:80	**C4**
3	**HL5** (6-Me)	45	98:2	*fac*-**C5**
4	**HL6** (4-Mes)	48	25:75	**C6**
5	**HL7** (5-Mes)	45	4:96	*mer*-**C7**
6	**HL8** (6-Mes)			

[a]Isolated yield. [b]*fac*/*mer* ratio determined by ^{1}H NMR.

Concerning the bulkier Mes group, no influence on the final configuration was again observed when substituted at position 4 (ligand **L6**) due to its outward directionality, obtaining the corresponding facial and meridional **C6** complexes as a statistical 1 to 3 mixture (entry 4). Nevertheless, the presence of the Mes group at position 5 leads to very crowded environments, notably in a facial configuration. As a result, ligand coordination yielded almost exclusively *mer*-**C7** complex (entry 5). Unfortunately, coordination of ligand **L8** with the Mes group adjacent to the pyridinic N atom was unsuccessful, even upon increasing the reaction temperature or preforming the carbene before adding the iron source (entry 6). It is worth noting that similar substitutions at the 5-membered NHC moiety in related Fe^{II} Py-NHC complexes have been described.[36] However, the smaller angle formed between the Mes group and the Fe−N bond due to the 6-membered pyridine ring is likely to increase the steric hindrance to such an extent that coordination is completely inhibited.

To evaluate the excellent selectivity toward *fac* and *mer* geometries promoted by **L5** (6-Me) and **L7** (5-Mes), we subsequently examined the impact of the NHC electronic nature. Thus, besides imidazole-based (Im) precursors, two additional ones were prepared for each substitution pattern from benzimidazole (BIm) and 4-methylimidazole (4-MeIm), which would afford more π-acceptor and more σ-donor NHC units, respectively. The synthesis of the NHC-varying ligand precursors **HL9**−**HL12** is depicted in Scheme 3.

Using the previously described complexation conditions (see Table 1), target **C9**−**C13** complexes were obtained in moderate yields (40−49%) (Table 2). As it can be observed, electronic properties of the NHC unit did not affect either the facial coordination selectivity with a 6-Me substitution (entries 1−3) or the meridional selectivity in the case of the 5-Mes substitution (entries 4−6), attaining even a complete selectivity for both configurations with the facial or meridional congeners remaining at best below the detection limit in ^{1}H NMR. Therefore, these results nicely indicated that the final configuration of the system is mostly determined by the azine substitution pattern in py-NHC complexes.

The remarkable and unanticipated facial selectivity exerted by the 6-Me substituted ligands **L5**, **L9**, and **L10** spurred us on to investigate the exact role played by this group. As a result, a new series of ligand precursors was prepared with varying functionalities at this position, including a more sterically hindered isopropyl (^{i}Pr) chain (**HL13**), several halogenated substituents (**HL14**−**HL16**), and a dimethylamine (NMe_2) group (**HL17**) (Figure 2). As shown in Scheme S3 (see the Supporting Information (SI)), **HL13**−**HL17** were prepared following rather similar synthetic protocols, as previously

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Scheme 3. Synthesis of Ligands HL9−HL12

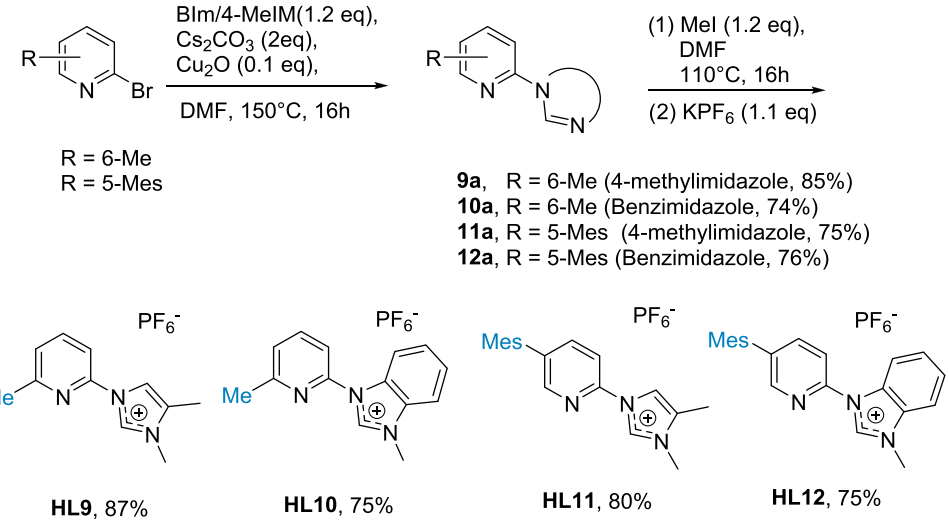
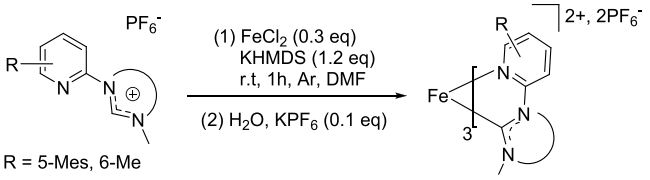


Table 2. Effect of the Nature of the NHC Unit on the Fe^{II} Complex Formation. Data from HL5 and HL7 are Presented for Comparison Reasons



entry	ligand precursor	yield (%)[a]	*fac*/*mer* ratio[b]	complex
1	**HL5** (6-Me, Im)	45	98:2	*fac*-**C5**
2	**HL9** (6-Me, 4-MeIm)	48	100:0	*fac*-**C9**
3	**HL10** (6-Me, bIm)	43	100:0	*fac*-**C10**
4	**HL7** (5-Mes, Im)	45	4:96	*mer*-**C7**
5	**HL11** (5-Mes, 4-MeIm)	49	0:100	*mer*-**C11**
6	**HL12** (5-Mes, bIm)	40	0:100	*mer*-**C12**

[a]Isolated yield. [b]*fac*/*mer* ratio determined by ^{1}H NMR.

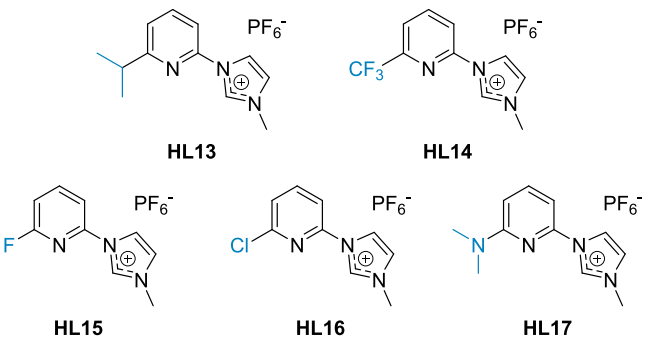
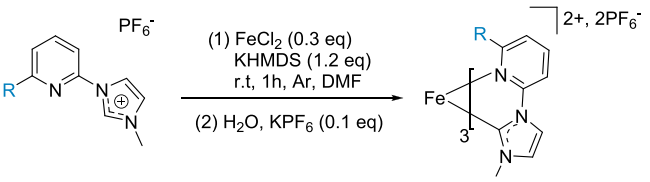


Figure 2. Series of ligand precursors with varying 6-substituted pyridine moieties.

described, starting from the appropriately substituted 2-halopyridine. In the case of **HL13** and **HL17**, previous introduction of their respective ^{i}Pr and NMe_2 groups was necessary.

The complexation results of precursors **HL13**−**HL17** are gathered in Table 3. The introduction of an ^{i}Pr group provided the same effect as a methyl group, *i.e.*, an exclusive facial isomerism, affording *fac*-**C13** in 30% yield (entry 1). The

Table 3. Coordination with Different Substituents at Position 6 on the Pyridine Moiety



entry	ligand precursor	yield (%)[a]	*fac*/*mer* ratio[b]	complex
1	**HL13** (R = ^{i}Pr)	40	100:0	*fac*-**C13**
2	**HL14** (R = CF_3)			
3	**HL15** (R = F)			
4	**HL16** (R = Cl)	18	100:0	*fac*-**C16**
5	**HL17** (R = NMe_2)			

[a]Isolated yield. [b]*fac*/*mer* ratio determined by ^{1}H NMR, with ≥1% as the detection limit.

electron-withdrawing halogenated substituents had more differing results (entries 2−4). Indeed, while no complex was obtained with **L14** (6-CF_3) and **L15** (6-F) attributable to an excessively reduced basicity of the pyridine, **L16** (6-Cl) led exclusively to *fac*-**C16** complex though in rather a poor yield (18%). This latter example is, however, of particular interest since it would allow for complex functionalization via cross-couplings or S_N^{Ar} reactions, for instance. In the case of the electron-donating NMe_2 group, no complex was isolated (entry 5). A tentative explanation for this could be a competitive coordination of the amino group to the metal ion that would probably inhibit the coordination of the NHC moiety. In consequence, from these results, we can infer that facial selectivity is not necessarily led by an agnostic interaction between the hydrogens of the methyl group and the pyridine ring since selectivity was not compromised with −Cl and −^{i}Pr substituents.

Given that the 6−CH_3 substituent could provide a similar steric hindrance as an internal −CH− of a phenyl group (Figure 3), we decided to examine the effect of extending the π-conjugation at the azine moiety as well.

Ligand precursors derived from quinoline (**HL18**) and quinoxaline (**HL19**) were thus synthesized (Scheme 4). **HL18** was prepared in 75% yield in a 2-step synthesis by reacting 2-

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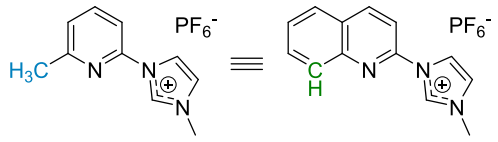


Figure 3. Steric Hindrance of CH_3 Compared to Internal CH.

chloroquinoline with *in situ*-generated potassium imidazolate, followed by quaternization with MeI. The higher reactivity of 2-chloroquinoxaline toward a S_N^{Ar} reaction allowed **HL19** to be directly obtained in 77% yield upon reaction with *N*-methylimidazole. The complexation of both π-extended H**L18** and **HL19** precursors was carried out under the same aforementioned conditions using $FeCl_2$ and KHMDS in DMF at r.t. (Scheme 4). Gratifyingly, the results were consistent with our hypothesis since the coordination afforded exclusively *fac*-**C18** and *fac*-**C19** complexes in 37 and 34% yield, respectively. As in the case of 6-Me (*fac*-**C5**) pyridine-based ligands, these results are in stark contrast with the stereochemistry obtained with other related ligands such as 2-(pyridin-2-yl)quinoline, which lead to the exclusive formation of the meridional isomer.[40]

X-RAY ANALYSIS

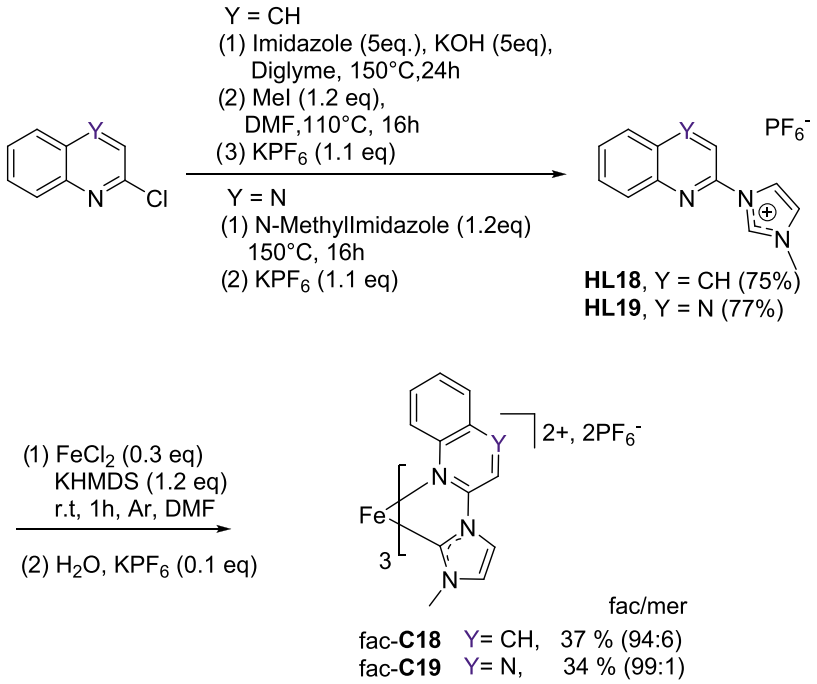
Suitable crystals of *fac*-**C5**, *mer*-**C7**, and *fac*-**C18** for X-ray diffraction were grown by slow evaporation of the corresponding methanol (*fac*-**C5**) or acetonitrile (*mer*-**C7** and *fac*-**C18**) solutions (see Table S1 and Figures SS90−S92). Figure 4 gathers the structures for the three complexes, displaying a distorted octahedral geometry and confirming the anticipated configurations. Some selected structural parameters are collected in Table 4.

Both structures of *fac*-**C5** and *fac*-**C18** show three bidentate ligands with the same orientation, resulting in a facial arrangement with each Fe−C bond trans to a Fe−N bond. In comparison to the computationally optimized structure of *fac*-**C0**,[33] average Fe−C bond distances (1.904 Å for *fac*-**C5** and 1.899 Å for *fac*-**C19** vs 1.948 Å for *fac*-**C0**) are significantly shorter, while average Fe−N bond distances are only slightly shorter (2.141 Å for *fac*-**C5** and 2.138 Å for *fac*-**C18** vs 2.072 Å for *fac*-**C0**). Nevertheless, the steric hindrance exerted by the imine unit is more evident when analyzing the conformation of the bidentate units, which exhibit marked tilting angles of 4.4−13.1 and 5.1−10.4° for *fac*-**C5** for *fac*-**C18**, respectively. Thus, it would be tempting to correlate the enhanced structural trans effect in *fac*-**C5** and *fac*-**C18** with respect to *fac*-**C0** with a higher kinetic trans effect, which would explain the obtained selectivity, but knowledge of the complexation mechanism is required.[12]

Furthermore, it is worth noting the presence of several short contacts in these facial complexes. These interactions involve mostly the coordinated atoms ($C_{carbene}$ or N_{azine}) to the iron center with a neighboring CH_3 group (d + 0.2 Å < sum of the van der Waals radii) for both *fac*-**C5** and *fac*-**C18**, or $N_{quinoline}\cdots H-C_{quinoline}$[8] (d + 0.4 Å <sum of the van der Waals radii) in the case of *fac*-**C18**. Since $C_{carbene}\cdots Cl$ intramolecular interactions have been reported,[42] this type of short contacts might be also envisaged for *fac*-**C16**. As a result, it could be likely that these interactions play a non-negligible role in the here-obtained facial selectivity.

The structure of complex *mer*-**C7** shows, on the contrary, a meridional arrangement where only two ligands are oriented in the same direction, resulting in a C_1 symmetry. While the mutually trans Fe−C and Fe−N bonds are rather comparable to those in the facial complexes, more remarkable differences are observed for the other four metal−ligand bonds. In fact, the mutually trans Fe−C bonds are longer (∼1.945 Å) and the mutually trans Fe−N bonds are shorter (∼1.981 Å) than those in *fac*-**C5** and *fac*-**C18**, being consistent with the stronger trans effect of a carbene ligand relative to an azine moiety. Moreover, the less crowded coordination sphere allows the ligands to adopt almost coplanar conformations (tilting angles = 0.20−

Scheme 4. Synthesis of Ligand Precursors HL18 and HL19 and Synthesis of Complexes *fac*-C18 and *fac*-C19



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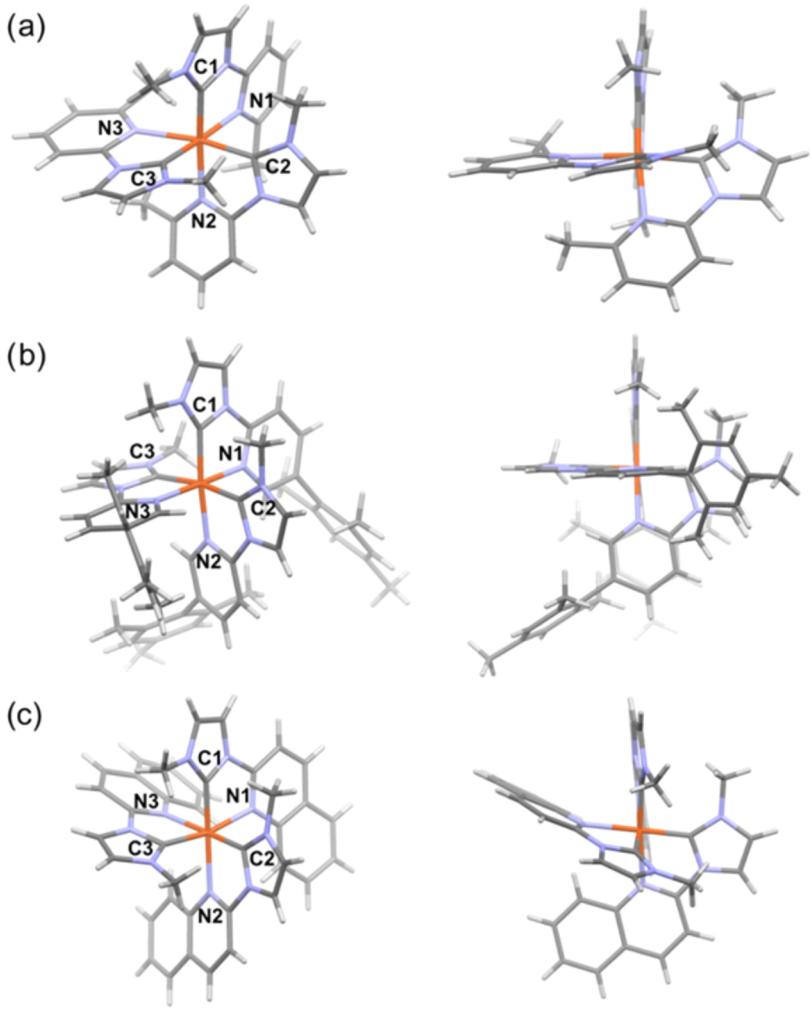

Figure 4. Left View from the (pseudo)-C_3 axis and (right) side view of the X-ray crystal structures of complexes (a) *fac*-**C5**, (b) *mer*-**C7**, and (c) *fac*-**C18** with a partial labeling. Periphery atoms on the ligands are omitted for more clarity. CCDC for *fac*-**C5**: 2126499; *mer*-**C7**: 2126506; and *fac*-**C18**: 2126500.

Table 4. Selected Structural and Deformation Parameters for Complexes *fac*-C5, *fac*-C18, and *mer*-C7 (esd in Parenthesis)[a]

	fac-**C0**	*fac*-**C5**	*fac*-**C18**	*mer*-**C0**	*mer*-**C7**
Fe-N1 (Å)	2.073	2.105(3)	2.137(2)	2.038	1.975(3)
Fe-N2 (Å)	2.072	2.131(2)	2.135(2)	2.082	2.015(2)
Fe-N3 (Å)	2.070	2.188(2)	2.142(2)	2.042	1.988(3)
Fe-C1 (Å)	1.949	1.893(3)	1.907(2)	1.942	1.903(3)
Fe-C2 (Å)	1.950	1.910(3)	1.899(2)	1.992	1.934(3)
Fe-C3 (Å)	1.947	1.910(3)	1.892(2)	1.983	1.956(3)
C1-Fe-N2 (°)	172.6	174.1(1)	168.6(9)	173.2	173.4(1)
C2-Fe-N3 (°)	172.8	169.9(1)	173.6(9)		
C3-Fe-N1 (°)	172.7	176.6(1)	176.1(9)		
C2-Fe-C3 (°)				170.0	169.9(1)
N1-Fe-N3 (°)				177.0	177.8(1)
$\langle Fe - X \rangle$ (Å)[b]	2.010	2.023	2.019	2.013	1.962
ζ (Å)[c]	0.368	0.710	0.717	0.245	0.185
Σ (°)[d]	64.9	81.4	85.9	67.7	61.2
Θ (°)[e]	205.1	241.2	257.0	213.3	196.4

[a]The Deformation Parameters were Calculated with OctaDist.[41] For the Sake of Comparison, Data for DFT-Optimized *fac*-**C0** and *mer*-**C0** are Provided as Well, with Partial Labeling Analogous *fac*-**C5** and *mer*-**C7**, Respectively[23] [b]Average metal−ligand bond. [c]Length distortion calculated as $\xi = \sum_{i=1}^{6} (d_i - \langle \mathrm{Fe} - X \rangle)$, with d_i being a metal−ligand bond. [d]Angle distortion calculated as $\Sigma \sum_{i=1}^{12} |90 - \phi_i|$ with ϕ_i being a *cis* ligand−metal−ligand angle. [e]Trigonal distortion calculated as $\Theta = \sum_{i=1}^{24} |60 - \theta_i|$, with θ_i being a torsional angle between the ligand atoms on opposite triangular faces of the octahedron viewed along the pseudo-threefold axis.

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2.38°), resulting in fewer intramolecular short contacts mainly consisting of L···H−$C_{pyridine}$[6] (L = $C_{carbene}$ or N_{azine}).

A more global comparison can be further done upon calculation of the stretching (ζ), angular (Σ), and trigonal (Θ) distortions from the ideal octahedral structure (Table 4).[41] Taking as reference the calculated structures of *fac*-**C0**, it is evident that the deformation was induced by the sterically demanding azine unit, with an increased asymmetry in bond lengths (ζ values) and worse M−L interaction (Σ and Θ values, with $\Sigma = 0°$ and $\Theta = 0°$ for a perfect octahedron). Interestingly, complex *mer*-**C7** displays the most regular coordination sphere even when compared with parent *mer*-**C0**.

THEORETICAL CALCULATIONS

The structural and energetic differences between the *fac* and *mer* isomers for some of the complexes synthesized in this work have been studied with quantum chemistry methodologies to better understand the observed isomer preferences. Geometry optimizations were performed with the density functional theory (DFT) method in acetonitrile.[43−46] Final energies, computed on top of the optimized geometries, were determined with the post-Hartree−Fock method domain-based local pair natural orbital coupled-cluster singles, doubles, and perturbative triples [DLPNO-CCSD(T)].[47,48] The latter method provides energies close to the full CCSD(T) procedure,[49] the gold standard in quantum chemistry, prohibitive for large-size molecules like the ones synthesized in this work. Solvent effects for ACN or DMF provide almost indistinguishable results (Table S2). Full computational details are included in the SI.

The *fac* vs *mer* thermodynamic stabilities are shown in Figure 5. It becomes apparent that the *fac* disposition is the most stable configuration for most of the Fe^{II}-NHC family, a trend that globally agrees with the *fac* preference observed experimentally. However, more specific analyses reveal that the thermodynamic stabilities of the isomers are not the only factors dictating the *fac*/*mer* preferences, suggesting that other factors also play an important role. These factors would include complexation kinetics and/or intramolecular interactions, as previously suggested.

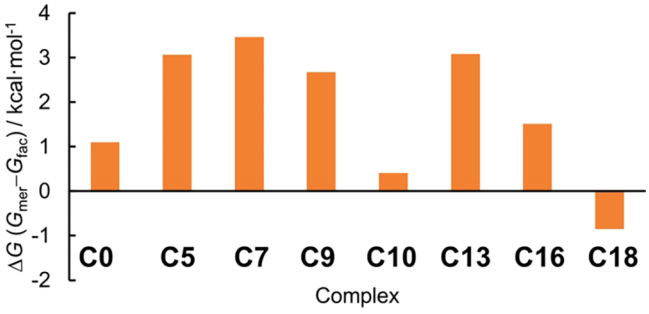

Figure 5. Gibbs energy difference (ΔG) between the *fac* and *mer* isomers for some selected complexes synthesized in this work. Optimized geometries and vibrational frequencies computed at the DFT/B3LYP-D3 level of theory. Energies refined with the DLPNO-CCSD(T) method in DMF.

The Gibbs energy differences between the two isomers are not equal for all complexes (Figure 5). Whereas **C5**, **C7**, **C9**, and **C13** clearly favor the *fac* disposition (ΔG ~3 kcal/mol), complexes **C0**, **C10**, and **C18** show almost degenerated energies for both *fac* and *mer* arrangements (ΔG ~ ±1 kcal/mol). This is coherent with the *fac*/*mer* isomer distribution (6.6:93.4) observed for **C0**[23] and the *fac* preference (>98:2) exhibited by **C5**, **C9**, and **C13**. An eye-chasing case is **C7**, synthesized mostly as a *mer* isomer, although the *fac* disposition is thermodynamically more stable. This evidences that the complexation mechanism penalizes the *fac* isomer. Meanwhile, the ΔG values for **C10** and **C18** also do not explain the *fac*/*mer* ratio higher than 94:6 as observed experimentally in terms of thermodynamic stabilities, again pointing to extrinsic factors in the complexation process.

A particularly interesting comparison is that of complexes **C0** and **C5**. Both compounds only differ in the methyl group at the 6 position of the pyridine ring in **C5**, which is key to inducing opposite isomer preferences: **C0** is mostly *mer*, whereas **C5** is *fac*. A detailed analysis of the Fe−N and Fe−C connections, as well as the noncovalent interactions between the methyl groups and the C and N dative atoms coordinated to iron, is presented in Figure 6 and Table 5. For *fac*-**C5**, the distances clearly indicate that the methylation at the 6 position of the pyridine enlarges the three Fe−N distances 0.134 Å (on average) due to steric congestion of the three methyl groups. This effect is illustrated in Figure 6b. As a result, on the opposite side of the molecule, the Fe−C and the Im-CH_3···C distances decrease by −0.005 and −0.051 Å, respectively (Table 5). Regarding *mer*-**C5**, the structural deformation is similar. On average, the Fe−N distances increase by 0.114 Å, and the Fe−C distances decrease by −0.004 Å, respectively. The largest difference between the *mer* isomers of **C0** and **C5** is found for the Im-CH_3···C and Im-CH_3···N distances, which decrease by−0.072 Å.

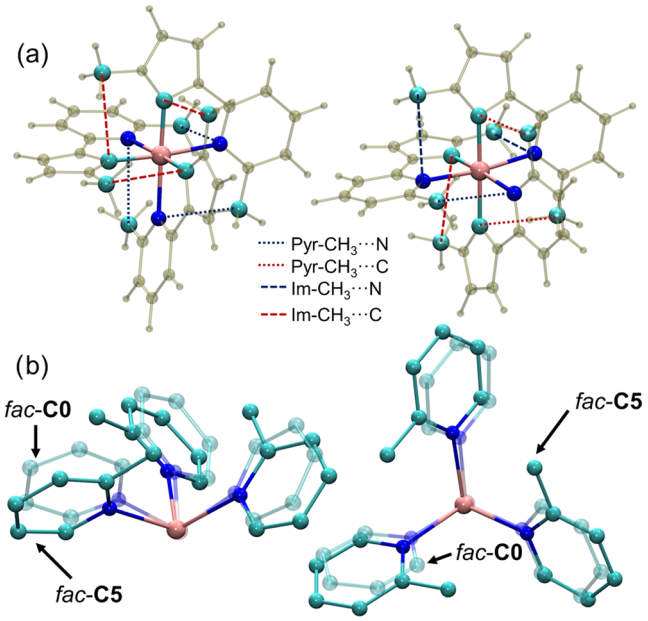

Figure 6. (a) DFT/B3LYP-D3-optimized structures of the *fac* (left) and *mer* (right) isomers for **C5** and nomenclature of the intramolecular noncovalent interactions. (b) Side (left) and top (right) views of the *fac* isomer of **C0** (transparent texture) and **C5** (solid texture). For the sake of clarity, only the Fe centers and the pyridine rings are shown.

Globally, the net structural deformations upon methylation from **C0** to **C5** are almost equal for both isomers, even though the deformation is not equally distributed into the same

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Table 5. Relevant Interatomic Distances (in Å) for the Optimized Structures of *fac* and *mer* Isomers of C0 and C5 as Defined in Figure 6

bond/interaction	*fac*-**C0**	*fac*-**C5**	*mer*-**C0**	*mer*-**C5**
Fe−N	2.057	2.195	2.065	2.126
	2.057	2.190	2.030	2.200
	2.057	2.189	2.025	2.135
Fe−C	1.931	1.926	1.927	1.977
	1.930	1.926	1.968	1.917
	1.931	1.926	1.974	1.963
Im-CH_3···C	3.175	3.117	3.205	3.166
	3.174	3.135		
	3.172	3.117		
Im-CH_3···N			3.224	3.136
			3.209	3.121
Pyr-CH_3···C				3.047
				3.005
Pyr-CH_3···N		3.045		3.023
		3.043		
		3.039		

molecular parameters. These subtle—yet important—differences impact the relative stability of both isomers. As a matter of fact, the thermodynamic stability of *fac*-**C5** is about ∼2 kcal/mol larger than that of *fac*-**C0** with respect to the respective *mer* counterparts (Figure 5). This difference is the combination of diverse attractive and repulsive interactions. In this regard, the more pronounced stretching of the more labile Fe−N bonds (as compared to Fe−C) in *fac*-**C5**, as well as the more favorable Pyr-CH_3···N noncovalent interactions (as compared to the Pyr-CH_3···C), in which the methyl group can act as a Lewis base[50] and the N as a Lewis acid (N is more electronegative than C and thus more acidic), may explain the slight energetic differences.

The noncovalent interactions available only in *fac* arrangements also explain the stability of the *fac*-**C7** structure with respect to the *mer* isomer. The *fac* arrangement clearly shows triple T-stacking interactions between the methyl groups of the mesityl substituent and the aromatic rings (Figure 7), whereas, in the *mer* disposition, only two 5-Mes groups interact due to intrinsic reasons. It can be hypothesized that the reason to obtain **C7** mainly as a *mer* isomer (Table 2) is likely due to steric congestion in the complexation mechanism. This analysis evidences again the need to consider mechanistic and other extrinsic factors besides the intrinsic stability of the isomers to fully understand the *fac*/*mer* isomerism in this kind of Fe^{II}-NHC complexes.

GROUND-STATE CHARACTERIZATION

Optical and electrochemical properties of all facial and meridional complexes were investigated by UV−vis spectroscopy (Figure 8) and cyclic voltammetry (see Figures S93−S102), and the main results are collected in Table 6. As in previously reported compounds,[5] the UV−vis spectra of all complexes present three main absorption bands. The intense bands below 300 nm correspond to $^1(\pi \rightarrow \pi^*)$ transitions centered on the ligands. At longer wavelengths, two distinct broader and less intense MLCT bands are found at longer wavelengths corresponding to Fe−carbene $^1(d \rightarrow \pi^*_{NHC})$ transitions (320−420 nm) and Fe−azine $^1(d \rightarrow \pi^*_{azine})$ transitions, which extend well into the visible region (380−650 nm).[5] Nevertheless, a clear effect of the ligand configuration can be observed (Figure 8, left and middle). On the one hand, the transitions are less intense in facial complexes due to symmetry reasons.[23] Moreover, NHC modification can result in different modulations of the $^1(d \rightarrow \pi^*_{azine})$ bands. While more donor NHC moieties correlate well with lower-energy transitions in facial complexes, with MLCT energies varying as *fac*-**C10** (bIm) > *fac*-**C5** (Im) > *fac*-**C9** (4-MeIm) (Figure 8, left), a no apparent trend is obtained in the meridional series (MLCT band energy: *mer*-**C7** (Im) > *mer*-**C11** (4-MeIm) > *mer*-**C12** (bIm)) (Figure 8, middle) (*vide infra*). Indeed, BIm carbene units lead to a blue shift in facial coordination (*fac*-**C10** vs *fac*-**C5**) but cause a red shift in meridional coordination (*mer*-**C12** vs *mer*-**C7**).

The effect of the 6-substitution on the pyridine unit appears to be related to the relatively elongated 6-Me-pyridine−Fe interaction (Figure 8, right). Actually, *fac*-**C5** (6R = Me) and *fac*-**C16** (6R = Cl) show almost identical spectra in spite of

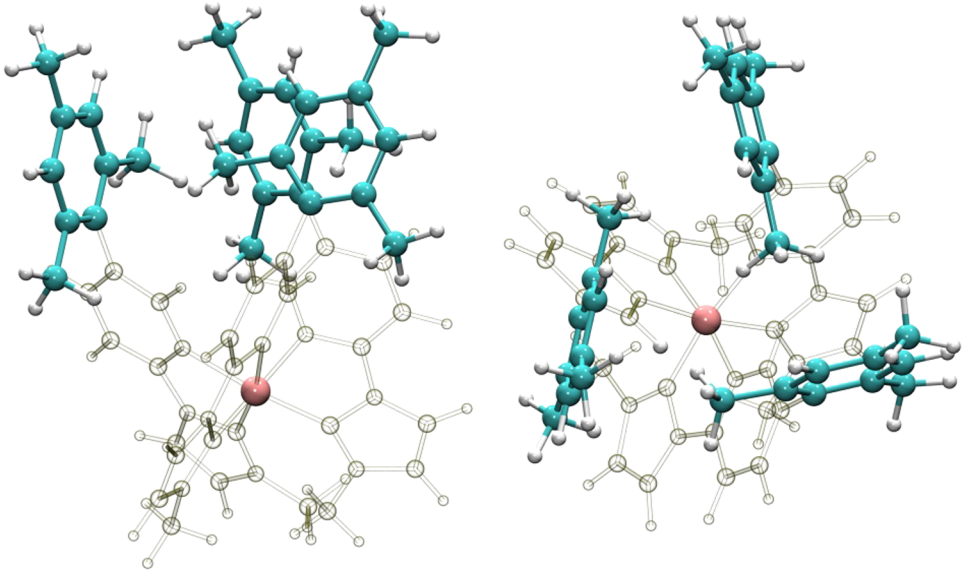

Figure 7. Side (left) and top (right) views of the *fac*-**C7** structure. 5-Mes groups are highlighted as cyan (carbon) and white (hydrogen) spheres.

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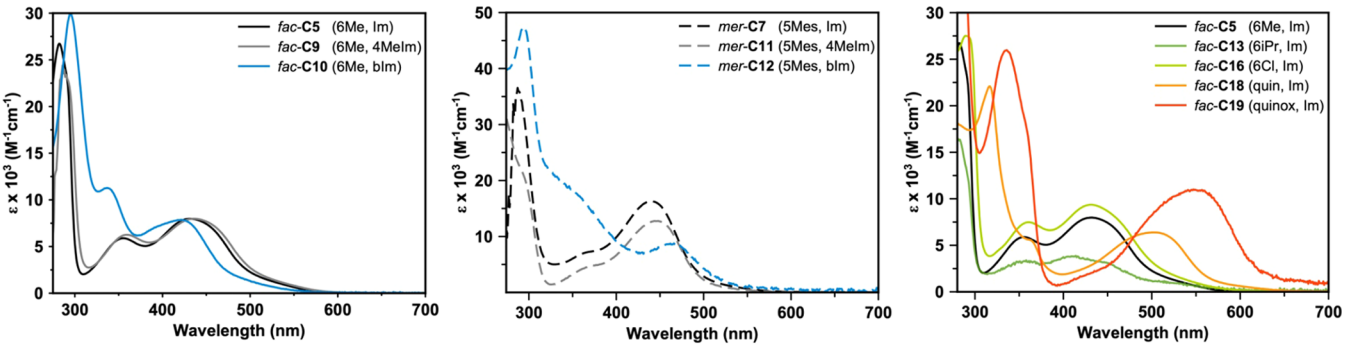

Figure 8. UV−vis spectra in air-equilibrated acetonitrile solution at room temperature to show: (left) NHC influence in facial series; (middle) NHC influence in meridional series; and (right) azine variation through substitution or π-conjugation extension.

Table 6. Photophysical and Electrochemical Data for the Facial or Meridional Fe^{II} Complexes

complex	$\lambda_{abs\text{-}max}$ (nm) [ε($M^{-1}.cm^{-1}$)][a]	E_{ox}[b] [V/SCE]	E_{red} [V/SCE]	ΔE [V][c]
fac-**C5**	285 [26 795]	0.69 (rev)	−1.59 (qr)	2.28
	355 [5891]		−1.87 (rev)	
	435 [7957]			
fac-**C9**	286 [24 398]	0.62 (rev)	−1.59 (rev)	2.21
	357 [6237]		−1.90 (qr)	
	436 [7928]			
fac-**C10**	293 [29 471]	0.95 (rev)	−1.19 (rev)	2.14
	340 [11 185]		−1.60 (rev)	
	419 [7839]			
fac-**C13**	283 [16 359]	0.81 (rev)	−1.42 (rev)	2.23
	351 [3274]		−1.91 (irr)	
	416 [3810]			
fac-**C16**	278 [25 731]	0.86 (rev)	−1.59 (irr)	2.45
	293 [27 383]			
	362 [7470]			
	428 [9332]			
fac-**C18**	317 [22 077]	0.79 (rev)	−1.41 (irr)	2.20
	367 [5157]		−1.63 (irr)	
	508 [6354]		−1.75 (irr)	
fac-**C19**	338 [25 764]	1.08 (qr)	−0.82 (irr)	1.90
	428 [1916]		−1.16 (irr)	
	555 [10 856]		−1.51 (irr)	
mer-**C7**	289 [35 856]	0.73 (rev)	−1.86 (rev)	2.59
	362 [6987]			
	444 [16 187]			
mer-**C11**	296 [19 460]	0.66 (rev)	−1.90 (rev)	2.56
	368 [4589]			
	451 [12 640]			
mer-**C12**	293 [47 267]	1.06 (rev)	−1.37 (rev)	2.43
	360 [16 852]		−1.84 (irr)	
	464 [8562]			

[a]Measured in air-equilibrated CH_3CN solution at 25°C. [b]First oxidation potential. Potentials are quoted *vs* SCE. Recorded in CH_3CN using NBu_4PF_6 (0.1M) as supporting electrolyte at 100 mV s^{-1}; under these conditions, $E_{1/2\ (Fc+/Fc)}$ = 0.39V/SCE; rev = reversible; qr = quasi-reversible; irr = irreversible. [c]Electrochemical band gap ($\Delta E = E_{ox} - E_{red1}$).

their distinct electronic nature. In the case of *fac*-**C13** (6R = ^{i}Pr), the less intense UV−vis spectrum might be attributed to the increased steric hindrance of ^{i}Pr in comparison with Me or Cl. In stark contrast, benzannulation of the azine unit does induce a noticeable change in the absorbance spectrum, evidencing the increased accepting character with bands red-shifted *ca.* 3200 cm^{-1} for *fac*-**C18** (quinolyl) and *ca.* 5000 cm^{-1} for *fac*-**C19** (quinoxalyl) with respect to *fac*-**C5** (pyridyl) (Figure 8, right).

Concerning the redox properties, cyclic voltammetry of the complexes was carried out with SCE as the standard electrode and ferrocene as internal reference (Fc^{+}/Fc = 0.39V/SCE in acetonitrile). At positive potentials, these bidentate *fac*/*mer* complexes display reversible Fe^{II} to Fe^{III} oxidations in most cases, with potentials ranging between 0.62 and 1.08 V (Table 6). These values can be nicely correlated to the π-back donation of Fe^{II} to both NHC and azine moieties, which results in higher reduction potentials due to the stabilization of the t_{2g}-like orbital (HOMO). For instance, the increasing

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electron-withdrawing character on going from 4-MeIm to Im to bIm is reflected on an anodic shift of the oxidation potentials in both facial and meridional series (*fac*-**C9** < *fac*-**C5** < *fac*-**C10** and *mer*-**C11** < *mer*-**C7** < *mer*-**C12**). The same effect is observed when the azine unit is modified, *e.g.*, 0.69 V for *fac*-**C5** (pyridyl), 0.79 V for *fac*-**C18** (quinolyl), and 1.08 V for *fac*-**C19** (quinoxalyl).

Noticeable differences in the cathodic events ascribed to the reduction of the ligands are obtained as well (−1.91 to −0.82 V). In comparison to **C0** (−1.94 V, −1.97 V), reduction is likely to take place at the azine unit for *mer*-**C7** (−1.86 V) and *mer*-**C11** (−1.90 V), together with *fac*-**C18** and *fac*-**C19**, which show drastic cathodic shifts that correspond well with their increased acceptor character and thus, lower π^* orbitals. Moreover, the introduction of more π-conjugated NHC units (bIm) seems to contribute to more delocalized LUMO, resulting in lower potentials as well (*fac*-**C10** vs *fac*-**C5** and *fac*-**C9**, *mer*-**C12** vs *mer*-**C7**, and *mer*-**C11**).[5] However, the interpretation is less straightforward for the other complexes (*fac*-**C13** and *fac*-**C16**). In fact, the steric hindrance exerted by the azine units in facial configurations affects their coordination to the metallic center and, thus, the exact nature of the Fe−L interactions.

In consequence, these results highlight the subtle interplay between electronic effects and geometry. For facial pyridyl-based complexes, MLCT energies appear to mainly depend on the carbene donor strength (Figure 8, left and right). However, meridional complexes are more sensitive to the π-conjugation of the carbene unit, which would explain the opposite effect observed for bIm-bearing complexes *fac*-**C10** (Figure 8, left) and *mer*-**C12** (Figure 8, middle). In the latter case, the higher stabilization of the t_{2g}-like orbital as a result of the π-accepting character of the bIm moiety is largely offset by the concomitant stabilization of the π^* orbital, resulting in a red-shifted MLCT band. On the other hand, selection of π-extended azines as in *fac*-**C18** and *fac*-**C19** leads to a remarkable stabilization of the ligand π^* orbitals, nicely agreeing with the noticeable red shift of their corresponding UV−vis spectra.

CONCLUSIONS

In this contribution, we have described some ligand design guidelines to selectively access both facial and meridional isomers in azine-NHC Fe(II) complexes by means of a room-temperature protocol. As we have demonstrated, the origin of this outstanding selectivity stems from the bulkiness of the azine unit in nearby positions to the N coordinating atom. Interestingly, an increase of the steric hindrance at the α position (6-substituted derivatives) with relatively small substituents such as a methyl group or a chlorine atom resulted in a surprising *fac* configuration. In fact, considering the common orientation of these substituents, similar substitution patterns of bidentated ligands have been typically reported to lead to the *mer* isomer. Moreover, π-extended quinoline and quinoxaline moieties exerted a similar preference for facial isomerism. In the case of the highly sterically demanding mesitylene group, target complexes were only achieved upon moving away the steric hindrance to position 5 of the azine unit, where *mer* isomers were mostly obtained instead. In spite of some preliminary computational calculations, an explanation of the *fac*/*mer* selectivity requires a deeper investigation on the complexation mechanism, which is currently ongoing in our laboratory.

Structural characterization of these complexes has revealed that *fac* isomers possess more distorted coordination spheres than the *mer* isomers, which could be deleterious for their EE dynamics. Nevertheless, it is also reasonable to think that the presence of more interligand interactions in *fac* derivatives could impede Fe−N bond elongation, which has been shown to be the main coordinate driving the relaxation.[33] On the other hand, the nature of the NHC unit and the extension of π-conjugation appeared to be the key parameters controlling their optical and redox properties, with both isomers exhibiting distinct behaviors. The use of π-extended azines is of particular relevance due to their low-lying MLCT manifold, which results in an improved absorbance in the visible range. In addition, the possible reduction of the electronic coupling with the metal-centered states could contribute to longer-lived MLCT. However, a proper analysis of the *fac*/*mer* influence on the EE kinetics will be the object of a following contribution.

Thus, this work highlights the great potential of this type of complexes that capitalize on the synergetic effect of combining a certain substitution pattern at the azine unit together with NHC. As a result, nonlabile iron(II) complexes can not only be prepared with a ligand-encoded stereochemistry, but also with optoelectronic properties that can be easily fine-tuned as a function of both the azine and the NHC moieties.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at https://pubs.acs.org/doi/10.1021/acsorginorgau.2c00038.

NMR, spectroscopic and electrochemical data for complexes, and additional computational and photophysical details (PDF)

Accession Codes

CCDC 2126499−2126500 and 2126506 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Author Contributions

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Notes

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