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# Low-temperature synthesis and characterization of lead-free $\text{BaTi}_{0.89}\text{Sn}_{0.11}\text{O}_3$ piezoelectric powders

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and Zdravko Kutnjak<sup>6</sup> and Mimoun El Marssi<sup>2</sup>

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## Abstract:

Barium stannate titanate ( $\text{BaTi}_{0.89}\text{Sn}_{0.11}\text{O}_3$ ,  $\text{BTS}_{11}$ ) is one of the reported lead-free materials that exhibits high dielectric and piezoelectric properties near ambient temperature. In this work,  $\text{BTS}_{11}$  powders were elaborated via a soft chemistry method at very low temperatures using two synthesized approaches combining sol-gel and hydrothermal processes. In the first approach, the effect of the barium precursor on the purity and the crystallinity of the  $\text{BTS}_{11}$  powders was studied, while in the second approach, the influence of the hydrothermal temperature on the  $\text{BTS}_{11}$  properties was analyzed. Pure and uniform fine powders were obtained at 180°C, the appropriate mass fraction of  $\text{BTS}_{11}$  which leads to high suspension stability in water is below 3%. These eco-friendly and pure piezoelectric powders can be used without further heat-treatment in several applications such as lead-free piezoelectric decontamination of organic pollutants and photocatalyst, hydrogen generation and as fillers in flexible nanocomposites nano-generators for energy harvesting and energy storage.

Key words: Lead-free piezoelectric, barium stannate titanate, sol-gel, hydrothermal, low-temperature chemistry.

## 1. Introduction :

Lead-based perovskite materials have been considered the best candidate for energy storage, energy harvesting, piezoelectric photo-degradation and hydrogen production applications, owing to their unrivalled dielectric properties. However, due to the lead toxicity, and the strict limitation of its usage, replacing it with an eco-friendly material became an urgent and legitimate task. For this purpose, barium titanate ( $\text{BaTiO}_3$ , abbreviated as BT) is regarded as one of the potentially promising lead-free ceramics alternatives. However, due to its low dielectric constant, high curie temperature (120 °C) and narrow temperature stability, researchers have enhanced its properties by site-doping, hence many barium titanate systems have been studied over the years, such as BCZT [1], BST [2], BCTS [3],[4], and finally BTS or barium stannate titanate[5].

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1 It was reported that replacing titanium ions with tin ions shifts the Curie temperature ( $T_c$ ) toward lower values.  
2 When Tin content is about 11%, a quasi-quadruple point (with coexisting cubic, tetragonal, orthorhombic and  
3 rhombohedral phases) that exhibits double morphotropic phase boundaries coexist together near the ambient  
4 temperature, leading to high dielectric constant, high piezoelectric coefficient, and therefore enhanced energy  
5 storage[5]. Moreover, Yong gang Yao et al. obtained the highest piezoelectric coefficient  $d_{33}$  of 697pC/N,  
6 five times higher than that of pure BaTiO<sub>3</sub>, in BaTi<sub>0.89</sub>Sn<sub>0.11</sub>O<sub>3</sub> in the vicinity of two converged triple points with  
7 coexisting multi-phase [6]. Meanwhile, Kalyani et al. showed that Sn substitution at low concentration  
8 drastically increased the piezoelectric coefficient  $d_{33}$  from 190 to 425 pC/N comparable to that exhibited by lead-  
9 zirconate titanate (PZT) [7]. Barium stannate titanate was mainly studied in the bulk ceramics elaborated by a  
10 conventional solid-state reaction [5],[8],[9], or in the thin films using a sol-gel solution [9,10]. However, very  
11 few researchers highlighted BTS powders elaborated at low temperatures and their potential applications as  
12 fillers in organic polymers to enhance the energy storage and the energy harvesting in nano-generators for  
13 energy harvesting and motion-sensing [10], or as BTS<sub>11</sub> nanoparticles for high dyes degradation and hydrogen  
14 generation efficiency [11].

15 The present work provides a detailed study on structural and physicochemical properties of (BaTi<sub>0.89</sub>Sn<sub>0.11</sub>O<sub>3</sub>,  
16 BTS<sub>11</sub>) powders synthesized by a combined Sol-Gel/Hydrothermal method at very low temperatures using two  
17 approaches which combine the advantages of excellent homogeneity, high purity, and low reaction temperature  
18 over other synthesis methods.

## 19 **2. Experimental details:**

### 20 **2.1. Chemicals:**

21 In order to prepare the BTS<sub>11</sub> ceramics, barium acetate Ba(CH<sub>3</sub>COO)<sub>2</sub> 99%, barium chloride (BaCl<sub>2</sub>) 99.99%,  
22 barium hydroxide (Ba(OH)<sub>2</sub>) 98.8%, titanium isopropoxide (C<sub>12</sub>H<sub>28</sub>O<sub>4</sub>Ti) 97%, and tin chloride (SnCl<sub>2</sub>.2H<sub>2</sub>O)  
23 98%, were used as raw materials, whereas acetic acid and 2-Methoxyethanol were used as the solvents. All  
24 reagents and solvents of chemical grade were used without any further purification.

### 25 **2.2. Preparation of BTS<sub>11</sub> powders**

26 The BaTi<sub>0.89</sub>Sn<sub>0.11</sub>O<sub>3</sub> (here after called as BTS<sub>11</sub>) powders were elaborated according to two synthesized  
27 approaches. The first one involves the study of the influence of three barium precursors on the purity and  
28 microstructure of the BTS<sub>11</sub> powders since it is mostly the primary source of impurities in low-temperature  
29 synthesis, and in the hydrothermal process more specifically [12],[13]. The three considered barium precursors  
30 have the distilled water as the common solvent which instantly leads to the titanium precursor (titanium  
31 isopropoxide) hydrolysis, thus, the necessity to separate the A site (Barium) from the B site (Tin and Titanium)  
32 of the perovskite as described in the first approach's experimental section. The second approach highlighted the  
33 effect of the hydrothermal temperature ranging from room temperature to 180 °C by using the adequate barium  
34 precursor (barium acetate) selected in the first approach. In order to reach a better homogeneity and to avoid any  
35 intermediates products, all precursors were mixed together in a separate sol-gel step to form a BTS<sub>11</sub> precursor in  
36 this approach. The concentration of the NaOH was fixed to 10M in both approaches.

37 The NaOH concentration has been previously studied by Xiang Ji et al. in BCZT perovskite powders,  
38 and they deduced that the higher NaOH concentration, the higher the crystallinity and the better

39 homogeneity, they also found that the grain size decreases with the increase of the NaOH  
40 concentration [14]. Furthermore, Z. hannani et al. also used NaOH 10M to elaborate BCZT perovskite  
41 powders and achieved high cristallinity [1].

#### 42 **2.1a First approach: effect of barium precursor type**

43 An appropriate amount of titanium isopropoxide and tin chloride were mixed in isopropanol under an inert  
44 atmosphere (N<sub>2</sub>) and stirred continuously until a homogenous mixture was achieved in one hour. Drops of  
45 distilled water were added to obtain the gel precursor. The gel precursor was then dried overnight at 80 °C. Next,  
46 the obtained precursor powder was homogeneously dispersed in 70 ml of 10M NaOH solution and mixed with a  
47 solution of barium acetate, barium chloride or barium hydroxide dissolved in 50ml of distilled water in a round-  
48 bottom flask with a magnetic stirrer for 3 hours under nitrogen flow. After that, the obtained suspension was  
49 poured in a Teflon stainless-steel autoclave at 180°C in an oven for 24h. Finally, the resulting BTS<sub>11</sub> powders  
50 were washed with distilled water and ethanol several times and dried at 80°C for 12 hours without of any  
51 calcinations.

#### 52 **2.1b Second approach: effect of hydrothermal temperature**

53 In the second approach, an appropriate amount of the adequate precursor (barium acetate) was dissolved, under  
54 stirring at 40°C for 1 hour, in acetic acid instead of water to avoid the hydrolysis of tin and titanium precursors.  
55 At the same time, in a beaker, a stoichiometric ratio of tin chloride and titanium isopropoxide were added to 2-  
56 propanol under continuous stirring and flow of nitrogen gas. The two solutions were then mixed in a 100ml flat  
57 bottom flask and left stirring for 2 hours; ammonia was then added until the solution turned clear by adjusting  
58 the pH at 7. The resulted gel was then dried overnight at 80 °C to form a white powder that was then dispersed in  
59 a 10M NaOH solution for 45 minutes, and then the suspension was transferred to a Teflon-lined stainless-steel  
60 autoclave, sealed and heated at 25, 80, 130 and 180 °C for 24 h. The main goal was to achieve a crystalline and  
61 pure perovskite powders at the lowest possible temperature, so we started at room temperature (25 °C) and  
62 gradually increased the temperature by 50 °C until we reached the purest crystalline ceramic powder at 180 °C.  
63 Finally, the resulting BTS<sub>11</sub> powders were washed with distilled water and ethanol several times and dried at 80  
64 °C for 12 hours without the need of any calcinations.

### 65 **3. Characterizations**

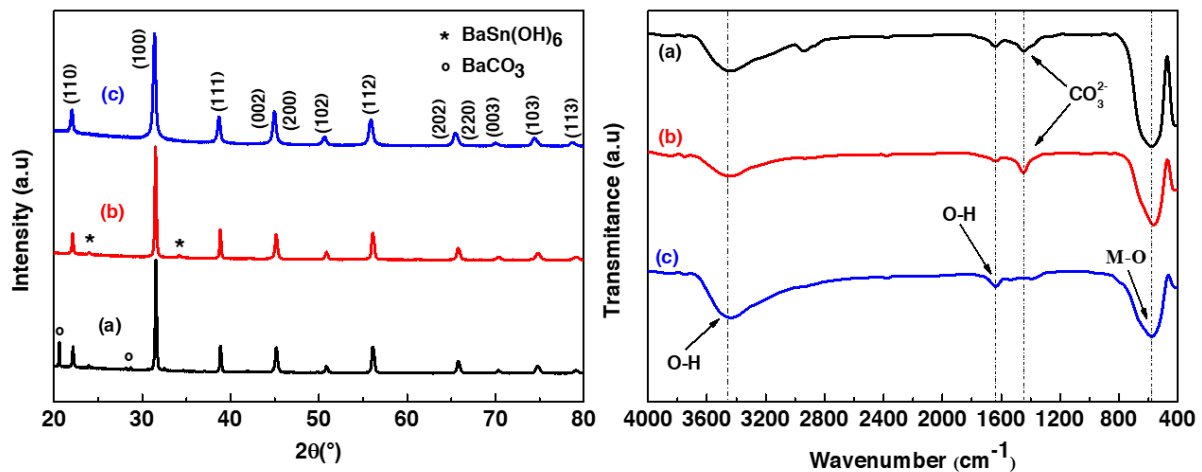
66 XRD measurements were carried out using an X-ray diffractometer brand "RIGAKU". X-rays were produced  
67 from a CuK $\alpha$  radiation source, having a wavelength equal to 1.54056 Å. X-ray diffraction spectra were recorded  
68 at room temperature in the range of 2 $\theta$  between 20° and 80° with angle steps of 0.02°. The scanning speed of  
69 5°/min was employed to perform simple phase identification. The scanning electron microscopy SEM  
70 observations were performed to examine our powders' morphology and approximate grain sizeona“Tescan  
71 Vega3” microscope. Moreover, in order to have a quick insight on the purity of our samples, the FTIR  
72 measurements have been carried out using a Bruker Vertex 70 spectrophotometer in transmission mode, 0.001 g  
73 of the powders were dispersed in 0.099 g of KBr matrix and analyzed in the frequency range 400 - 4000 cm<sup>-1</sup>.  
74 The combined thermogravimetric and differential thermal analyses (TG-DTA, Sytram LABSYS Evo) were also  
75 performed at a heating rate of 5 °C/min from room temperature to 700 °C in air. Differential scanning calorimetry

76 (Perkin Elmer Jade DSC) was used to point out the different phase transitions in a temperature range of 0 to 90  
77 °C. A NanoSp confotec-MR520 was used in order to perform the Raman spectroscopy, the samples were  
78 analyzed using a laser with a wavelength of 532 nm in a 1200 line/mm network, and lastly, a NanoZs, Malvern  
79 instruments Ltd, 7.12 was used for our zeta potential measurements at T = 25 °C with the switch time at t = 60 s.

#### 80 4. Results and discussion:

##### 81 4.1. First approach: selection of barium precursor type

82 Fig.1 displays the XRD patterns of the  $BTS_{11}$  powders synthesized at 180 °C using the three Ba precursors  
83 according to the first approach; it is worth noting that all the three samples exhibit a perovskite structure with a  
84 notable difference in the purity. When using barium hydroxide as barium precursor, an intense peak around  
85  $2\theta=21^\circ$  corresponding to  $BaSn(OH)_6$  impurity according to ICDD 15-780 is observed[15]. The presence of  
86 barium carbonate  $BaCO_3$  as a secondary phase is observed when using barium chloride which could be attributed  
87 to the reaction between  $Ba^{2+}$ , and  $CO_2$  which originated from the decomposition of organic solvents and/or  
88 dissolved in the Ba precursor's solvent.  $BaCO_3$  impurity can also occur as a byproduct due to atmospheric  $CO_2$   
89 contamination and the high thermodynamic stability of  $BaCO_3$  at high pH in aqueous solution compared to  
90  $BaTiO_3$  [16]. Although, interestingly, no secondary phase was detected in the XRD patterns when the barium  
91 acetate was used as a precursor, which is confirmed by the FTIR spectra of the three samples presented in Fig.1,  
92 the vanishment of the peak around  $1400\text{ cm}^{-1}$  and  $1500\text{ cm}^{-1}$  corresponding to the carbonates is noticeable when  
93 the barium acetate was used. Hence the choice of the barium acetate as the barium precursor in the second  
94 approach, further characterizations were done to complete the study.



95

96 Fig.1 XRD patterns and FTIR spectra of the  $BTS_{11}$  powders synthesized at 180°C by the first approach (a) barium hydroxide (b)  
97 barium chloride (c) barium acetate.

##### 98 4.2. Second approach: selection of hydrothermal temperature

99 In order to elucidate the effect of the hydrothermal temperature on the properties of  $BTS_{11}$  powders, barium  
100 acetate was used as the Ba precursor.

101 Fig.2 reveals the XRD patterns of  $BTS_{11}$  powders elaborated according to the second approach at different  
102 temperatures (25 °C, 80°C, 130 °C and 180 °C) with a scanning angle from 20° to 80°. It was observed that all

103 BTS<sub>11</sub> powders exhibit a perovskite phase along with the one synthesized at room temperature. It is also worth  
104 noting that the presence of a very weak peak at 24° may correspond to BaCO<sub>3</sub> in the samples synthesized at 25  
105 °C, 80 °C and 130 °C, which is due to the reaction of Ba<sup>2+</sup> with CO<sub>2</sub> despite inert atmosphere usage. Selvaraj et  
106 al. [17] reported that BaTiO<sub>3</sub> (BT) nanopowders prepared by combining sol–gel and hydrothermal processes at  
107 various hydrothermal temperatures were found to be in crystalline nature along with impurities like barium  
108 carbonate (BaCO<sub>3</sub>) and the intensity of diffraction peaks corresponding to BaCO<sub>3</sub> decreased with increasing  
109 temperature. However, at 180 °C, no secondary phases are detected, therefore the structure is highly pure. This  
110 demonstrates that the precursor powders synthesized by Sol-Gel is entirely transformed into well-crystallized  
111 BTS<sub>11</sub> powder in the hydrothermal step which is also confirmed by the thermogravimetric analysis of the powder  
112 synthesized at 180 °C presented in the fig. 3-a. The total weight loss observed in temperature range of 30 °C -  
113 700 °C is very low, 7% approximately. The first weight loss is due to the dehydration phenomena. It occurs at a  
114 temperature range of 40-160 °C. The second weight loss observed at 160 to 230 °C is ascribed to the chemical  
115 dehydroxylation and the decomposition of solvents, and the last low weight loss after 580 °C is associated to the  
116 thermal decomposition of remaining organic compounds [18].

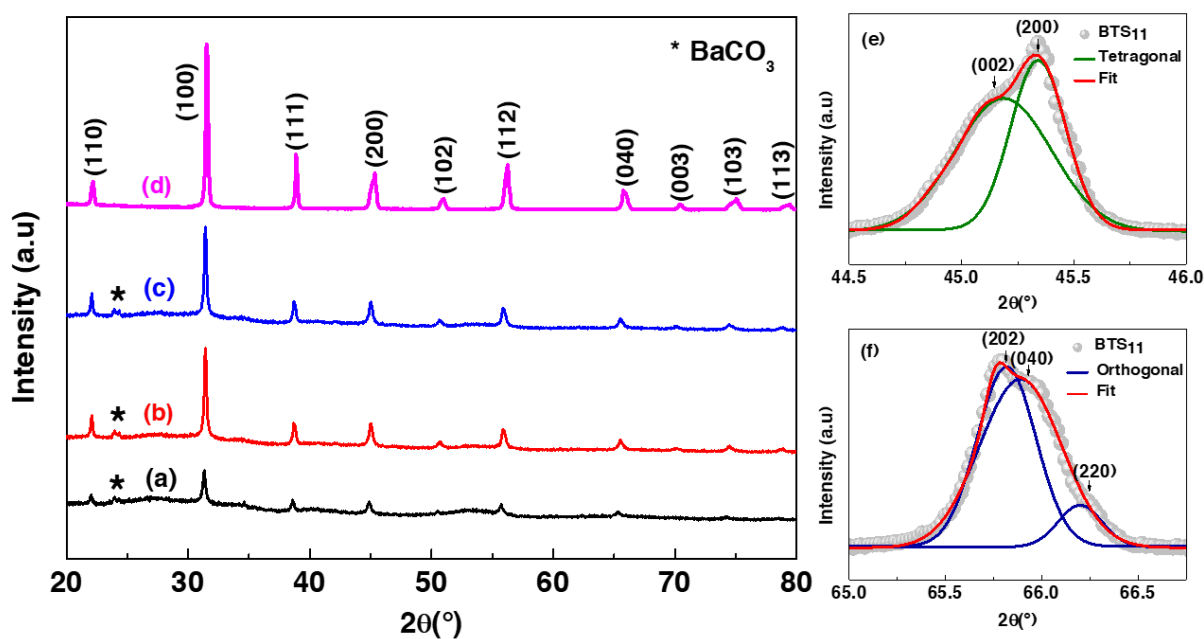
117 The FTIR spectra of BTS<sub>11</sub> powders synthesized at different temperatures were performed to study the purity of  
118 our powders more deeply as presented in fig.4. When the reaction temperatures are 25 °C and 80 °C, a weak  
119 absorption peak corresponding to CO<sub>3</sub><sup>2-</sup> at around 862 cm<sup>-1</sup> can be attributed to the few impurities of CO<sub>3</sub><sup>2-</sup> that  
120 were also detected by XRD. With the temperature increase, the extremely weak absorption peak of CO<sub>3</sub><sup>2-</sup>  
121 disappears. Moreover, the absorption peaks at 3000–3600 cm<sup>-1</sup> (the symmetric and asymmetric stretching  
122 vibrations of O–H), 1633 cm<sup>-1</sup> (the bending vibration of H–O–H) and 1421 cm<sup>-1</sup> (the symmetric stretching of  
123 C=O) also decline gradually, indicating the improvement of the chemical purity of BTS<sub>11</sub> powders. Finally,  
124 when the temperature approaches 180 °C, the intensity of the peaks corresponding to O–H and the C=O liaisons  
125 nearly disappear, indicating good purity of the obtained powders [19].

126 Multiphase such as rhombohedral (R,R3m), orthorhombic (O,Amm2), tetragonal (T, P4mm) and also cubic (C,  
127 Pm3m), were observed in the BaTi<sub>0.89</sub>Sn<sub>0.11</sub>O<sub>3</sub>, system by many researchers[11],[20],[21],[22]. The splitting  
128 observed in the enlarged view of the 45° XRD peak (fig.2-e) suggests the presence of a tetragonal phase, while  
129 the presence of a triplet at 65° (fig.2-f) indicates the existence of an orthogonal phase as well [23],[24],[25].  
130 Thus, to confirm these observations, full-pattern Rietveld refinement was performed and presented in fig.3. The  
131 presence of the tetragonal, orthorhombic and the rhombohedral was confirmed with a phase fraction of 55.77%,  
132 18.59% and 25.64% respectively. The refinement parameters are presented in table.1.

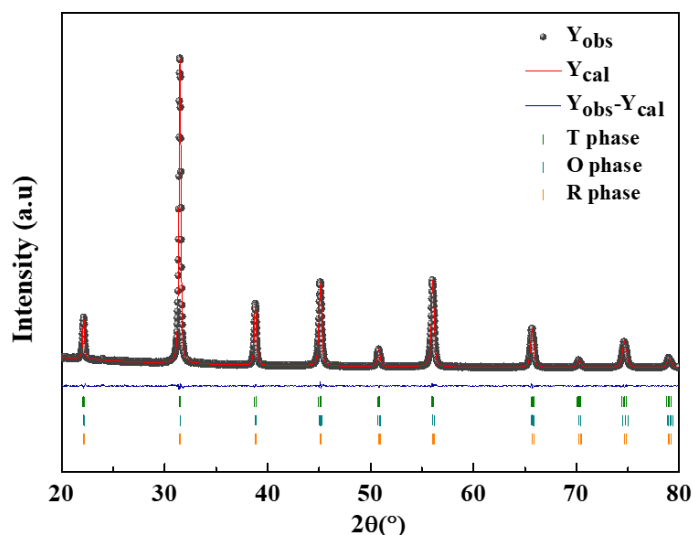
133 The Differential scanning calorimetry (DSC) was performed on the sample synthesized at 180 °C as presented in  
134 fig.4-b. The presence of only one weak endothermic peak in a temperature range close to ambient temperature,  
135 between 30°C and 45°C implies the existence of a multiphase [18]. It is well known that a sudden change in  
136 spontaneous polarization during a first-order transition will produce latent heat and specific heat dissipation at  
137 the Curie temperature, meanwhile, the tricritical behavior refers to a special thermodynamic condition in which a  
138 first-order transition becomes a second-order transition, so due to the discontinuity of polarization, it involves a  
139 discontinuous change in specific heat, hence the transition can almost be undetectable and the intensity of the  
140 peak is much smaller from the first-order transition [26],[27]. Thus, the presence of one endothermic peak in a  
141 temperature range between 0°C and 90 °C, can leads that T<sub>R-O</sub>, T<sub>O-T</sub> and T<sub>c</sub> are all overlapped in the temperature

142 range, between 30 °C and 45 °C. Therefore, high values of the dielectric constant and piezoelectric coefficient  
 143 can be achieved in this temperature range due to the properties of the multiple phase boundaries [11],[21].

144 Raman spectroscopy is a versatile method to probe structural properties at the nanoscale and is often used to  
 145 study structural transitions in ferroelectrics. Fig.4 shows Raman spectra at room temperature of  $BTS_{11}$  powders  
 146 synthesized at 25 °C, 80 °C, 130 °C and 180 °C. The modes at 190  $cm^{-1}$ , 253  $cm^{-1}$ , 305  $cm^{-1}$ , 513  $cm^{-1}$  and 716  
 147  $cm^{-1}$  corresponding to  $A_1(LO)$ ,  $A_1(TO)$ ,  $B_1, E(TO, LO)$ ,  $A_1(TO)$  and  $A_1, E(LO)$  respectively [28]. Raman peaks at  
 148 around 305 and 716  $cm^{-1}$ , characteristics bands of the ferroelectric phases, show weak peaks at low  
 149 temperatures, and become more intense and sharper with increasing the synthesis temperature. Likewise, the  
 150 intensity of the cubic paraelectric phase characterized by two peaks centered at 253 and 513  $cm^{-1}$  is weak at low  
 151 temperatures and well-defined at 180 °C, which also suggests the presence of multiphase. The Raman peak at  
 152 870  $cm^{-1}$  is due to the B site vacancies originating from the incomplete substitution of tin in the barium titanate  
 153 matrix. The peak disappears when the temperature exceeds 130 °C, suggesting a complete conversion of our raw  
 154 materials into the perovskite structure [29]. It is worthy to note that when the synthesizing temperature increases,  
 155 the characteristic peaks of the perovskite structure around 190, 305, 513 and 716  $cm^{-1}$  are getting sharper, which  
 156 implies that the crystalline quality of the  $BTS_{11}$  powder is improved, in perfect correlation with the XRD results  
 157 analysis[28][30].



158  
 159 Fig.2 XRD patterns of  $BTS_{11}$  powders synthesized at a) 25°C, b) 80°C, c) 130°C and d) 180°C and the enlarged views of 45° XRD  
 160 peak (e) and 65° peak (f) of the  $BTS_{11}$  powder synthesized at 180 °C according to the second approach.



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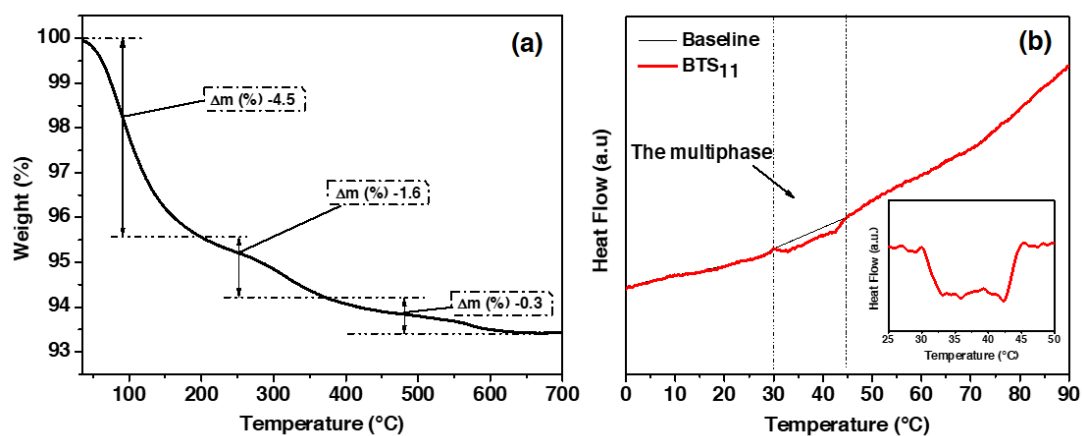
Figure 3 Full pattern Rietveld refinement of the  $BTS_{11}$  particles synthesized at 180 °C

163

**Table 1.** Refinement parameters of  $BTS_{11}$  by full pattern Rietveld refinements.

x	space group	a (Å)	b (Å)	c (Å)	$\alpha, \beta, \gamma$ (°)	Fraction (%)	GOF	$R_{exp}$	$R_{wp}$
0.11	<i>P4mm</i>	4.017506	4.017506	4.027725	90, 90, 90	55.77	1.4	5.49	7.73
	<i>Amm2</i>	4.017593	4.009575	4.023733	90, 90, 90	18.59			
	<i>R3m</i>	5.679052	5.679052	6.959726	90, 90, 120	25.64			

164

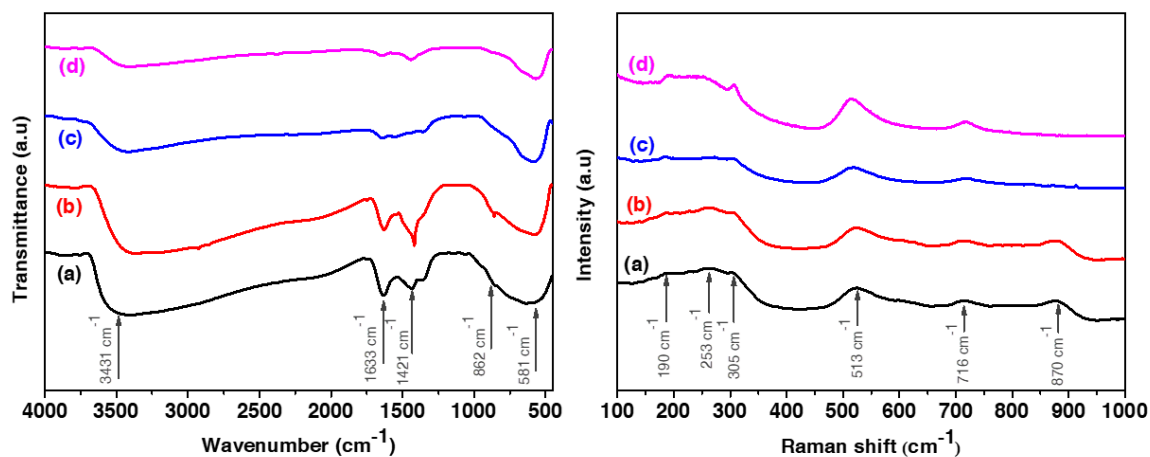


165

Fig.4 (a) ATG curve and (b) DSC heat flow curve of the  $BTS_{11}$  powder synthesized by the second approach at 180°C (The inset shows the zoom region from 25 to 50 °C).

167



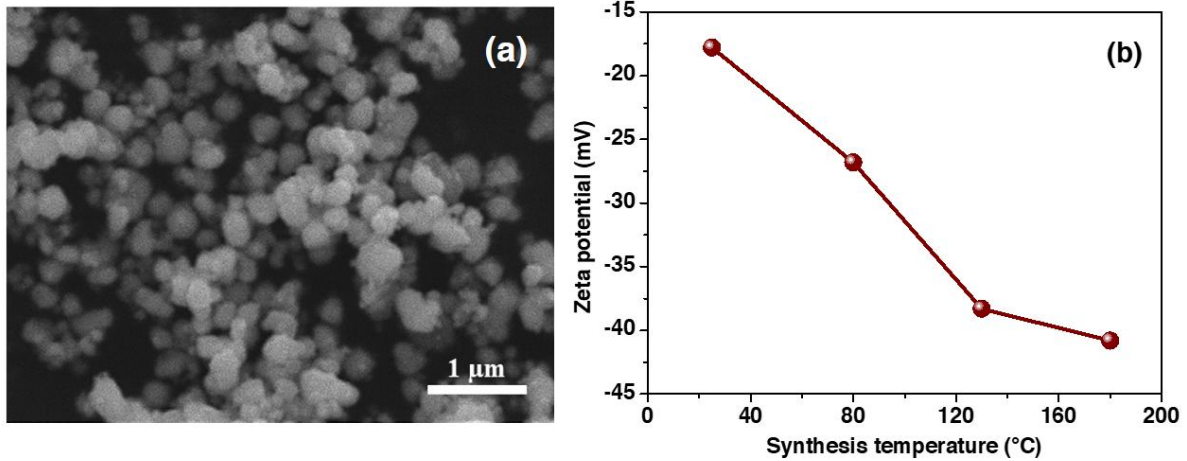


168

169 *Fig.5 Raman and FTIR spectra of the BTS<sub>11</sub> powders synthesized at a) 25 °C, b) 80 °C, c) 130 °C and d) 180 °C according to*  
 170 *the second approach.*

171 The SEM micrograph of the sample synthesized at 180 °C is shown in fig.5-a. The pure BTS<sub>11</sub> powder  
 172 synthesized at 180°C is spherically shaped and well homogenized with a distribution range of 475 nm.

173 To understand the surface charge of our powders and their electrostatic stabilization in a suspension during  
 174 colloidal processing, zeta potential measurements have been conducted in distilled water. The zeta potential of a  
 175 particle in a solution is determined by the charges on the various parts of the particle, the amount of associated  
 176 solvent and the ionic strength of the solution. The zeta potential is a measure of the strength of the intermolecular  
 177 forces between the particle and the solvent, and it is a measure of the repulsive tendency of the associated  
 178 solvent to the particle[31]. The larger the zeta potential, the stronger the associated intermolecular forces and the  
 179 more difficult it is for a particle to slip or migrate through a solution[32]. If all particles in a suspension have a  
 180 large negative or positive zeta potential ( $|\xi| > 30$  mV), they will tend to repel each other and the particles will not  
 181 clump together leading to a stable solution. However, if the particles have low zeta potential values ( $|\xi| < 30$   
 182 mV), the repulsive force is not sufficient to prevent the particles from clumping together and flocculating [33].  
 183 Thus, it is crucial to control the surface chemical interactions between BTS<sub>11</sub> particles and different additives,  
 184 such as electroactive polymers and organic pollutants in different media where aqueous suspension are  
 185 preferred owing to environmental considerations. The zeta potential of our as-prepared powders (1% in mass)  
 186 measured immediately after the colloidal preparation (Fig.5-b) shows negative surface charges and decreases  
 187 when the synthesis temperature increases starting from a value of -17.8 mV to -40.8 mV. However, when the  
 188 temperature is above 130 °C, the zeta potential is higher in absolute value than 40 mV and the system  
 189 is considered completely stable. These results agree with FTIR analysis showing the decrease of the number  
 190 of carbonate and hydroxyl groups remaining in the particles when the hydrothermal temperature increases.



191

192 *Fig.6 (a) SEM micrograph of the BTS<sub>11</sub> powder synthesized according to the second approach at 180 °C and (b) Synthesis*  
 193 *temperature dependent zeta potential.*

194 The percentage of the filler in the organic polymer matrix is a critical parameter in the functional properties of  
 195 the nanocomposites generators, such as the dielectric constant, breakdown strength and the energy storage  
 196 density[34],[35],[36]. It was reported that high mass fraction of the filler, leads to higher dielectric constant but  
 197 low breakdown strength[37].However, low fillers not only improves the breakdown strength of the  
 198 nanocomposites, but also remarkably enhance their mechanical strength [38]. Thus, low mass fraction is  
 199 advisable. Fig.6 represents the zeta potential of the BTS<sub>11</sub> powder in distilled water using different mass fraction  
 200 ranging from 1% to 5%to test the stability and the dispersibility of the particles. The result shows that above 3%,  
 201 the zeta potential starts drastically increasing showing that the system become unstable due to the agglomeration  
 202 of the particles[39], however when using a low filler fraction (less than 3%), the BTS<sub>11</sub> exhibits high stability and  
 203 dispersibility and no further surface modification is required. The highly negative zeta potential yielding to  
 204 repulsive forces at the surface and a high degree of dispersibility and stability is found when the BTS<sub>11</sub> mass  
 205 fraction is 3%.

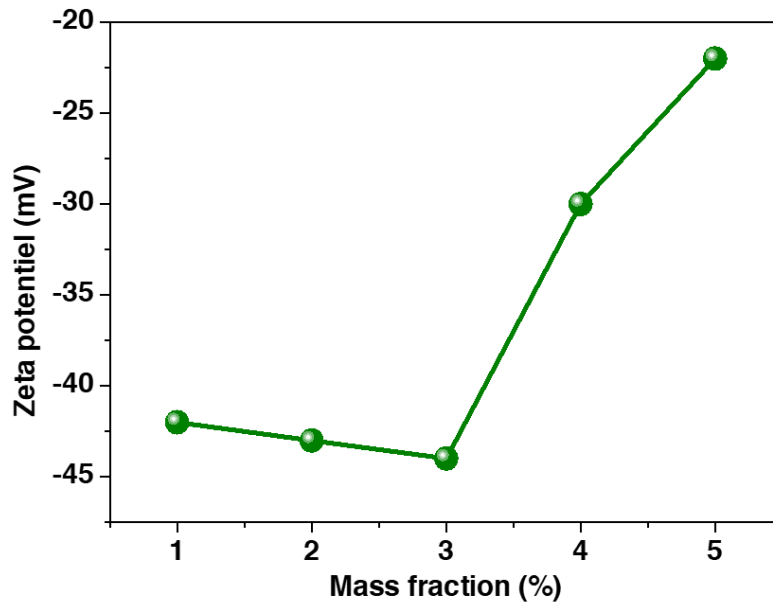


Fig.7 Mass fraction dependent Zeta potential of the  $BTS_{11}$  powder in distilled water synthesized according to the second approach.

So, it is worth noting that the  $BTS_{11}$  powder prepared at  $180\text{ }^{\circ}\text{C}$ , tend to be easily dispersed into water at low fraction and hence used as piezo-photocatalysis for dye decomposition and hydrogen production or as fillers in polymers nanocomposites for energy harvesting and energy storage. However, functionalizing the surface of the  $BTS_{11}$  particles is recommended when using a high filler fraction.

### 5. Conclusion:

Non-toxic and ecological  $BaTi_{0.89}Sn_{0.11}O_3$  ( $BTS_{11}$ ) powders were successfully prepared at very low temperature using sol-gel and hydrothermal methods combined without any further heat treatment. The effect of the barium precursor on the purity of the materials was studied. The hydrothermal synthesis temperature was proven to be a key factor for the crystalline purity of the powders and the improvement of their suspension stability. The mass fraction of 3% is the most suitable conditions for the  $BTS_{11}$  colloidal stability in the water. The Curie temperature of the powder was estimated to be around room temperature using the differential scanning calorimetry (DSC), which gives our eco-friendly powders huge advantages and potential applications as nanogenerators for energy storage, energy harvesting and hydrogen production and as piezoelectric photocatalysts.

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