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The Dole effect and its variations during the last 130,000 years as measured in the Vostok ice core

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Abstract. We review the current understanding of the Dole effect (the observed difference between the $\delta^{18}\text{O}$ of atmospheric O_2 and that of seawater) and its causes, extend the record of variations in the Dole effect back to 130 kyr before present using data on the $\delta^{18}\text{O}$ of O_2 obtained from studying the Vostok ice core (Sowers *et al.*, 1993), and discuss the significance of temporal variations. The Dole effect reflects oxygen isotope fractionation during photosynthesis, respiration, and hydrologic processes (evaporation, precipitation, and evapotranspiration). Our best prediction of the present-day Dole effect, +20.8 ‰, is considerably lower than the observed value, +23.5 ‰, and we discuss possible causes of this discrepancy. During the past 130 kyr, the Dole effect has been 0.05 ‰ lower than the present value, on average. The standard deviation of the Dole effect from the mean has been only ± 0.2 ‰, and the Dole effect is nearly unchanged between glacial maxima and interglacial periods. The small variability in the Dole effect suggests that relative rates of primary production in the land and marine realms have been relatively constant. Most periodic variability in the Dole effect is in the precession band, suggesting that changes in this global biogeochemical term reflects variations in low-latitude land hydrology and productivity or possibly variability in low-latitude oceanic productivity.

Introduction

Our understanding of the response of the biosphere to Pleistocene climate change is actually quite limited. For the terrestrial realm, our most detailed knowledge comes from extensive studies of pollen in sediments which record climate during and after the last glacial termination [e. g., COHMAP, 1988]. Pollen data for earlier times are extremely informative but geographically restricted. In the marine realm, the most ambitious attempts to understand the response of the biosphere to climate have focused on studies of primary productivity in the past inferred from two proxy indicators: organic carbon accumulation rates [e. g., Sarnthein *et al.*, 1988, 1992] and foraminiferal taxonomy [Mix, 1989a, b]. Paleoproductivity estimates made from many other proxies have enlivened the discussion with respect to specific areas of the ocean.

Proxy studies of the response of the terrestrial and marine biosphere are limited in that they do not generally provide information on a global scale, although with enough effort

they can potentially be made to do so. Ice core reconstructions of the concentrations of bioactive gases in air provide an integrated global signal which complements the proxy studies cited above. Variations in the concentrations of the most abundant bioactive gases (CO_2 , CH_4 , N_2O) reflect, in part, global scale changes in the marine and terrestrial biospheres. Atmospheric concentrations of CH_4 have been found to vary with a period corresponding to that of the precession of the equinoxes [Chappellaz *et al.*, 1990]. Concentrations are higher when temperatures are warmer and precipitation is greater. Glacial-interglacial variations in the concentration of CO_2 have been attributed to changes in the fertility and carbon export of the upper water ecosystem [e. g., Mix, 1989a; Knox and McElroy, 1984], although variations in ocean circulation and other factors must also have played an important role (e. g., Boyle, 1988; Broecker, 1989). On a shorter timescale, the transient minimum in the CO_2 concentration of air at the end of the last glacial termination reflects rapid growth of the land biosphere [Nefel *et al.*, 1988].

The $\delta^{18}\text{O}$ of O_2 is an additional variable which reflects the global responses of the land and marine biospheres to climate change, albeit in a complex manner. The Dole effect is defined as the difference between the $\delta^{18}\text{O}$ of atmospheric O_2 in air and the $\delta^{18}\text{O}$ of contemporaneous seawater. The magnitude of the Dole effect, which today is about +23.5 ‰ [Kroopnick and Craig, 1972], mainly reflects the isotopic composition of O_2 produced by marine and terrestrial photosynthesis, as well as the extent to which the heavy isotope is discriminated against during respiration. As previously discussed, $\delta^{18}\text{O}$ of

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atmospheric O₂ varied over glacial/interglacial timescales in response to changes in the δ¹⁸O of seawater [Horibe *et al.*, 1985; Bender *et al.*, 1985, Sowers *et al.*, 1993]. Other factors, which are of secondary importance in controlling the δ¹⁸O of atmospheric O₂, are (1) changes in terrestrial and marine fertility, (2) varying isotope fractionation associated with the hydrologic cycle, and (3) changes in respiratory isotope effects on either a species or community level.

The response time of δ¹⁸O of O₂ is comparable to the turnover time of O₂ in air with respect to photosynthesis and respiration, about 1.2 kyr (Table 1). In general, increases in the ratio of terrestrial to marine production will cause the δ¹⁸O of atmospheric O₂ to rise. Within the terrestrial realm, large-scale changes in ecology, precipitation, and humidity can all cause variability in the Dole effect.

In this paper we present a record of variations in the Dole effect over the last 130 kyr, based on the analysis of air occluded in the GISP2 and Vostok ice cores. We then explore the implications of this record for changes in the fertility of the land and marine biospheres and changes in hydrologic fractionation of oxygen isotopes during the Late Pleistocene.

Dole Effect Reconsidered

The δ¹⁸O of surface seawater today is, on average, close to zero on the SMOW scale, while that of O₂ is constant throughout the atmosphere at a value of about +23.5‰. Over the time period of interest here, photosynthesis and respiration are the most important reactions producing and consuming O₂. The isotopic composition of O₂ in air must therefore be understood in terms of isotope fractionation associated with these reactions [Lane and Dole, 1956; Berry, 1992; and references therein].

The δ¹⁸O of O₂ produced by photosynthesis is similar to that of the source water [e. g., Guy *et al.*, 1993]. The δ¹⁸O of O₂ produced by marine plants today is thus 0‰. The δ¹⁸O of O₂ produced on the continents has been estimated to lie between +4 and +8‰ [Dongmann, 1974; Farquhar *et al.*, 1993]. These elevated δ¹⁸O values are the result of elevated leaf water δ¹⁸O values resulting from evapotranspiration.

The δ¹⁸O of O₂ consumed by respiration is ~20‰ less than that of the source O₂ [Guy *et al.*, 1993; Kiddon *et al.*, 1993]. At isotopic steady state, the δ¹⁸O of O₂ molecules actually consumed by respiration must be the same as the δ¹⁸O of O₂ produced by photosynthesis [Lane and Dole, 1956]. Today this constraint is satisfied with δ¹⁸O of atmospheric O₂ = +23.5‰. In the event of deviations, the steady state value of the Dole effect would be restored with an e-folding time equal to the turnover time of O₂ in air, ~1.2 kyr.

We now recognize that several additional processes affect the δ¹⁸O of atmospheric O₂ [e. g., Berry, 1992]. The most important is evapotranspiration, which can cause very large enrichments in the δ¹⁸O of leaf water of land plants because of the preferential evaporation of the light isotope [Dongmann *et al.*, 1974; Farris and Strain, 1978; Forstel, 1978]. Due largely to the meticulous work of R. D. Guy and colleagues [Guy *et al.*, 1989, 1992], our estimates of isotope fractionation during photosynthesis and respiration are based on much better data than was available to Lane and Dole [1956]. In addition, we now have better estimates of the relative rates of land and ocean production. Finally, we can now begin to assess the impact of several additional processes on the Dole effect, including isotopic exchange between CO₂ and O₂ driven by photochemical processes in the stratosphere (first suggested by Dole *et al.* [1954], p. 66), mixing in the aphotic zone of the ocean and its influence on the δ¹⁸O of O₂ consumed by respiration in that environment, and equilibrium

Table 1a. Terrestrial Mass Balance of O₂ and δ¹⁸O of O₂

Production term		Production	Reference
Gross production excluding photorespired O ₂ (GPP)		14.1	Farquhar <i>et al.</i> [1993]
Gross production including photorespiration (14.1/0.69)		20.4	Farquhar <i>et al.</i> [1980]
Process	Fraction of respiratory O ₂ consumption	Isotope effect	Reference
δ ¹⁸ O of terrestrial photosynthetic O ₂ w. r. t. SMOW		4.4‰	Farquhar <i>et al.</i> [1993]
Discrimination against O ¹⁸ during respiration			
Dark respiration	59%	18.0	Guy <i>et al.</i> [1992, 1993]
Mehler reaction	10%	15.1	Guy <i>et al.</i> [1992]
Photorespiration	31%	21.2‰	Guy <i>et al.</i> [1992]
Flux weighted terrestrial respiratory isotope effect, excluding dark respiration		18.7‰	
Equilibrium enrichment in δ ¹⁸ O of leaf water w. r. t. air		+0.7‰	Benson and Krause [1984]
Terrestrial respiratory isotope effect (= 18.7‰-0.7‰)		18.0‰	
Terrestrial Dole effect		22.4‰	

Table 1b. Marine Mass Balance of O₂ and δ¹⁸O of O₂

Production term	Production (x10 ¹⁵ moles/yr)	Reference
Marine gross production (=4 x seasonal net production)	12	<i>Keeling and Shertz [1992]</i>
Marine gross production - recycling within the ocean	10.6	
Process	Isotope effect	Reference
δ ¹⁸ O of O ₂ produced by marine photosynthesis		
δ ¹⁸ O of marine photosynthetic O ₂	0 ‰	
Discrimination against ¹⁸ O during marine respiration		
Average marine ε _R (applies to euphotic respiration)	20 ‰	<i>Kiddon et al. [1993]</i>
Effective ε _R for respiration below the euphotic zone	12 ‰	<i>Bender [1990]</i>
Equilibr. enrichment in δ ¹⁸ O of O ₂ in water w.r.t. air	0.7 ‰	<i>Benson and Krause [1984]</i>
Effective ε _R for marine respiration*	18.9 ‰	
Marine Dole effect	18.9 ‰	

$$* = 0.95 \times (20 \text{ ‰} - 0.7 \text{ ‰}) + 0.05 \times 12 \text{ ‰}$$

isotope fractionation between O₂ in air and O₂ dissolved in seawater.

δ¹⁸O of O₂ Produced During Photosynthesis

In the terrestrial environment, gross production is the production term of interest relative to the Dole effect because all photosynthetic O₂ escapes from leaves to label the atmosphere before any is consumed intracellularly. This is the case because leaves are open to diffusion of CO₂, a gas whose atmospheric concentration is only 0.0016 times that of O₂.

In the ocean, photosynthesis takes place in the mixed layer or in adjacent underlying waters. Gas in the mixed layer equilibrates with the atmosphere over a timescale of weeks, while gas in the underlying seasonal thermocline equilibrates

over a timescale of months. We assume that the average timescale for gas equilibration between the euphotic zone and the atmosphere is 30 days, the euphotic zone is 60 m thick and its [O₂] is 250 μmol/L, the ratio of net/total production is 0.25, and net primary production is 8.5 mol m⁻² yr⁻¹ (Table 1). From these numbers we calculate that gross primary production is 34 moles m⁻² yr⁻¹. We also calculate that 88 % of photosynthetic O₂ goes into the atmosphere, and 12 % is consumed by respiration before it crosses the air/sea interface to label the atmospheric O₂ pool. We therefore discount gross oceanic production by 12 % in computing the δ¹⁸O balance of O₂.

Guy et al. [1993] recently showed that δ¹⁸O of O₂ produced by spinach thylakoids (photosynthetic membranes) was

Table 1c. Calculation of the global Dole effect, model leaf water enrichment, and atmospheric O₂ turnover time

	Isotope effect
Stratospheric diminution due to O isotope exchange of O ₂ and CO ₂	0.4 ‰
Calculated Dole effect*	20.8 ‰
Observed Dole effect	23.5 ‰
Leaf water enrichment giving observed Dole effect, all other terms constant	8.5 ‰
	Turnover time
Turnover time of atmospheric O ₂ **	1.2 kyr

*=[fraction of production on land x land Dole effect] + [fraction of production in oceans x marine Dole effect] - stratospheric diminution

**=(atmospheric inventory / [terrestrial production + marine production]) = (3.7x10¹⁹ moles / [20.4 + 10.6] x 10¹⁵ moles/yr)

indistinguishable from that of ambient water in which the thylakoids were suspended. Guy et al. and earlier workers [e. g., *Stevens et al.*, 1975] found slight enrichments ($\sim 1\text{‰}$) in the $\delta^{18}\text{O}$ of O_2 produced by phytoplankton relative to ambient waters. They attributed these enrichments to algal respiration in the photosynthesizing cultures rather than to photosynthetic fractionation. Thus the $\delta^{18}\text{O}$ of photosynthetic O_2 has the same isotopic composition as its source water, or very nearly so.

Marine photosynthesis. The $\delta^{18}\text{O}$ of surface seawater ranges from about -1‰ to $+1\text{‰}$ on the SMOW scale, averaging nearly zero [Craig and Gordon, 1965]. Thus the $\delta^{18}\text{O}$ of O_2 produced by marine photosynthesis is $\sim 0\text{‰}$.

Land photosynthesis. To understand the factors that control the $\delta^{18}\text{O}$ of terrestrially produced O_2 , one needs to understand the factors controlling the $\delta^{18}\text{O}$ of leaf H_2O . Leaf water ultimately originates from surface seawater, but its isotopic composition, relative to the source, is fractionated by equilibrium and kinetic effects attending evaporation and precipitation [Craig and Gordon, 1965]. The $\delta^{18}\text{O}$ of precipitation varies spatially and temporally and averages about -7‰ . This water is absorbed into the roots, stems, and branches of plants without significant fractionation. Water in leaves, however, is enriched in ^{18}O because, during evapotranspiration, the light isotope evaporates more rapidly. Quantitative descriptions of the leaf water enrichment assume a two-step process for evaporation. First, leaf water evaporates into saturated air within the "head space" (stomata) of the leaf with a composition fixed by equilibrium between leaf water and the head space vapor. Second, leaf water escapes from the stomata into free air which has a humidity $< 100\%$.

For a simple model in which evapotranspiration is simulated by the steady state evaporation of water from a sheet of paper, the "leaf water" enrichment is given by the "Craig equation:"

$$\delta_{\text{leaf}} - \delta_{\text{stem}} = \epsilon^* + (1 - h) \epsilon_k + h (\delta_{\text{atm}} - \delta_{\text{stem}}) \quad (1)$$

[Craig and Gordon, 1965; Dongmann, 1974; Farris and Strain, 1978], where δ is the $\delta^{18}\text{O}$ of water. The subscripts leaf, stem, and atm refer to leaf water, stem water (i. e., local meteoric source water), and atmospheric water, respectively. h is relative humidity, ϵ^* is the equilibrium isotope effect for evaporation of water (liquid) to water (gas), and ϵ_k is the kinetic isotope effect associated with the diffusion of water out of the stomata into the free atmosphere. ϵ^* is about $+9\text{‰}$ and ϵ_k is about 27‰ [Farquhar et al., 1993; Yakir, D., J. Berry, L. Giles, and C. B. Osmond, Isotopic heterogeneity of water in transpiring leaves: Identification of the component that controls the $\delta^{18}\text{O}$ of atmospheric O_2 and CO_2 , submitted to Plant and Cell Environment, 1994].

In the case where the relative humidity is 100% and atmospheric water is in equilibrium with soil water, the leaf water enrichment would be 0‰ relative to local meteoric water. In the case where the humidity is zero, the calculated leaf water enrichment is $\sim 36\text{‰}$ relative to local meteoric water. Other factors affecting the $\delta^{18}\text{O}$ of chloroplast water, and hence the $\delta^{18}\text{O}$ of photosynthetic O_2 , include the normal diurnal variation in humidity (which is at a minimum in the afternoon), elevation of temperatures inside leaves relative to

the surrounding air, non steady state conditions with respect to the leaf water balance (especially at night, when stomata close), variations in the kinetic isotope effect for the escape of leaf water due to the biological variability in the nature of the stomata [Farris and Strain, 1978], and variations in the $\delta^{18}\text{O}$ of H_2O gas in ambient air. Data in Figure 1, on the isotopic composition of leaf water from a forest in Julich, Germany, illustrate the expected diurnal cycle in the $\delta^{18}\text{O}$ of leaf water, together with variations demonstrating the existence of complex disequilibrium effects.

At any time when the humidity is $< 100\%$, the ideal value for the isotopic composition of leaf water can be modified by advection of water through the leaf to the site of evaporation

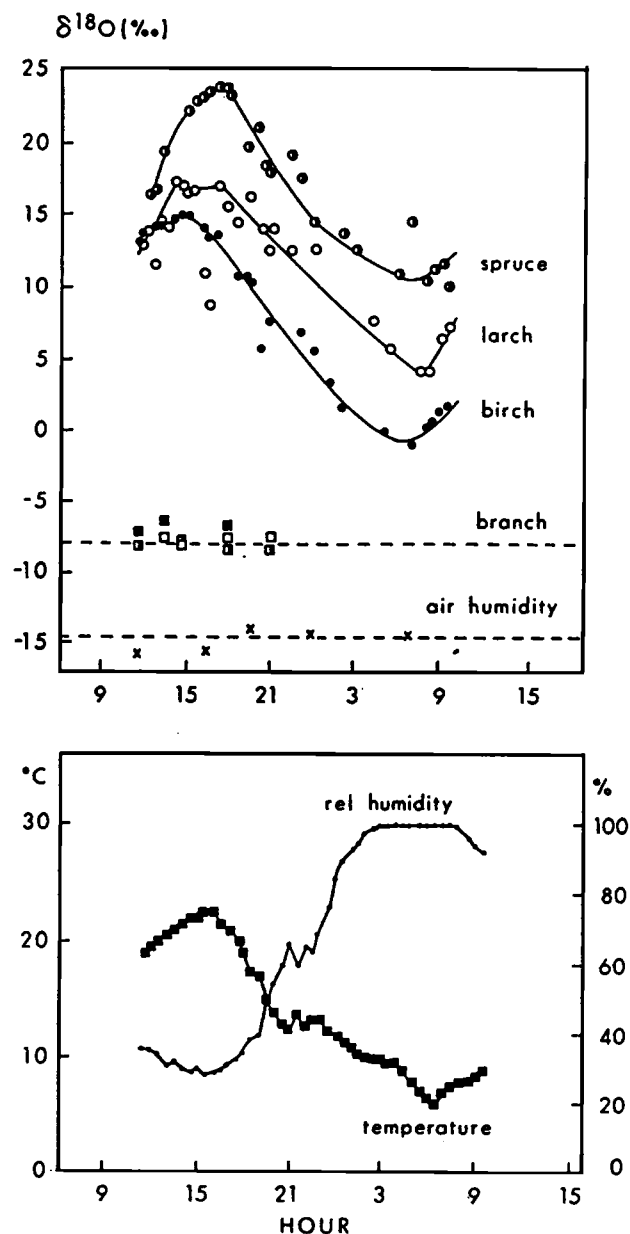


Figure 1. $\delta^{18}\text{O}$ of H_2O in leaf water of three trees, branch water, and air humidity as a function of time of day for a forest in Julich, Germany. Relative humidity and temperature are also shown. Redrawn from Forstel [1978].

and by back diffusion of leaf water from these sites. The isotopic composition of leaf water is thus intermediate between the $\delta^{18}\text{O}$ value predicted by the Craig equation and the $\delta^{18}\text{O}$ of stem water [Yakir *et al.*, 1989; Yakir, D., J. Berry, L. Giles, and C. B. Osmond, Isotopic heterogeneity of water in transpiring leaves: Identification of the component that controls the $\delta^{18}\text{O}$ of atmospheric O_2 and CO_2 , submitted to Plant and Cell Environment, (1994).; Farquhar *et al.*, 1993]. Of interest here is the ^{18}O enrichment in water in chloroplasts, where photosynthetic production of O_2 occurs. Farquhar *et al.* [1993] pointed out that this value might be governed by the Craig equation (equation (1)), since chloroplasts are near the sites of evaporation on the surface of leaves, and presented strong supporting evidence. Yakir *et al.* (Yakir, D., J. Berry, L. Giles, and C. B. Osmond, Isotopic heterogeneity of water in transpiring leaves: Identification of the component that controls the $\delta^{18}\text{O}$ of atmospheric O_2 and CO_2 , submitted to Plant and Cell Environment, 1994) presented data showing that the $\delta^{18}\text{O}$ of chloroplast water (and hence of photosynthetic O_2) was intermediate between that predicted by the Craig equation and that of stem water. Thus the basic control on the ^{18}O enrichment of water in chloroplasts is under debate.

With respect to the Dole effect, the value of interest is the globally averaged $\delta^{18}\text{O}$ of photosynthetic O_2 . This is equal to the time-averaged $\delta^{18}\text{O}$ of chloroplast water during the photoperiod, weighted for the rate of photosynthetic O_2 production. $\delta^{18}\text{O}$ of leaf water, and hence photosynthetic O_2 , varies over all scales of space and time, making the globally averaged value extremely difficult to estimate. Recently, Farquhar *et al.* [1993] attempted such a calculation in another context and estimated a value of + 4.4 ‰ for the productivity-weighted average $\delta^{18}\text{O}$ of leaf water and hence photosynthetic O_2 produced on land. We adopt this value in the calculations that follow but note that the + 4.4 ‰ value falls close to the lower end of the range of possible values suggested by Dongmann [1974].

$\delta^{18}\text{O}$ of O_2 Consumed by Respiration

Terrestrial respiration. Following Farquhar *et al.* [1980], we assume that the ratio f of photorespiration to carboxylation (or total carbon fixation) is given by equation (4) in their paper (p. 79):

$$\phi = (V_{\text{O}_{\text{max}}} / V_{\text{C}_{\text{max}}}) \times (\text{O}/\text{C}) \times (K_{\text{C}}/K_{\text{O}}) \times 10^{-3} \quad (2)$$

where $(V_{\text{O}_{\text{max}}} / V_{\text{C}_{\text{max}}})$ is the ratio of the maximum oxygenation velocity to the maximum carboxylation velocity of RuP₂ carboxylase-oxygenase, and O is the O_2 partial pressure (210,000 ppmV). C is the CO_2 partial pressure in plant cells in ppmV, assumed to be equal to 0.7 x partial pressure of CO_2 in air. $(K_{\text{C}}/K_{\text{O}})$ is the ratio of Michaelis - Menton constants for carboxylation and oxygenation, respectively (460/330). For the preindustrial value of 280 parts per million by volume (ppmv) CO_2 in air, $\phi = 0.31$. Following Guy *et al.* [1992], we assume that the Mehler reaction accounts for 10 % of terrestrial respiration. We assume that all dark respiration is via the cyanide-sensitive pathway. The partitioning of respiration is summarized in Table 1.

Respiratory oxygen isotope fractionation (ϵ_{R}) has recently been measured for the major biochemical pathways by Guy *et al.* [1989, 1992, 1993]. They determined ϵ_{R} to be about 18 ‰ for dark respiration by the cyanide-sensitive pathway and 25 ‰ for respiration by the alternative (cyanide-resistant) pathway. Guy *et al.* [1992] measured ϵ_{R} to be about 21.2 ‰ for photorespiration, and 15.3 ‰ for the Mehler reaction (in vivo photo-oxidation). We then calculate a value of $\epsilon_{\text{R}} = 18.8$ ‰ for terrestrial respiration (Table 1). The true value is likely to be slightly higher because of the influence of cyanide - resistant dark respiration.

Finally, we note that respiration consumes O_2 from the liquid pool of cells. The $\delta^{18}\text{O}$ in solution is enriched by 0.7 ‰ with respect to air [Benson and Krause, 1984], decreasing land fractionation with respect to the atmosphere by this amount.

Marine respiration. Kiddon *et al.* [1993] determined ϵ_{R} values for a range of microorganisms representative of common marine autotrophs and heterotrophs in culture. They measured ϵ_{R} values ranging from 15 ‰ to 25 ‰ and averaging close to + 20 ‰. Larger marine animals (as well as land plants and animals) have lower ϵ_{R} values due to diffusion limitation of fractionation. In these organisms, ϵ_{R} depends on rates at which light and heavy isotopes can diffuse to the site of respiration as well as on the biochemical fractionation [Lane and Dole, 1956; Epstein and Zeiri, 1988; Kiddon *et al.*, 1993]. However, these animals are not important in the global O_2 balance. The average value of ϵ_{R} for microorganisms studied by Kiddon *et al.* is close to that observed by Guy *et al.* [1992] and expected for cyanide-sensitive dark respiration based on work of Guy *et al.* [1989]. Nevertheless, for individual species there are differences which cannot be readily explained.

In this paper, we adopt a value of 20 ‰ for the marine value of ϵ_{R} . This value directly characterizes marine respiration in the euphotic zone of the ocean in and the upper part of the aphotic zone which equilibrates rapidly (i. e., seasonally) with the atmosphere. We neglect the fact that the marine respiratory isotope effect in nature must be higher in the light than in the dark (in which condition Kiddon *et al.* [1993] made all their measurements), since the isotope effect of 21.2 ‰ expected for photorespiration is higher than the isotope effect adopted here. Photorespiration is thought to be a minor pathway in the ocean. We note that respiration by the alternative pathway is associated with much more extensive isotopic fractionation than via the cyanide resistant pathway, but we have no way of assessing the relative rates for the two pathways. Finally, we note that the $\delta^{18}\text{O}$ of the dissolved O_2 pool of seawater is enriched with respect to the $\delta^{18}\text{O}$ of air by 0.7 ‰ [Kroopnick and Craig, 1972]. This enrichment reflects the fact that $^{18}\text{O}^{16}\text{O}$ is 0.7 ‰ more soluble in seawater than $^{16}\text{O}_2$ [Benson and Krause, 1984]. This enrichment decreases the ^{18}O fractionation relative to the atmosphere by 0.7 to 19.3 ‰.

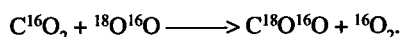
A small fraction of marine respiration occurs within more isolated parts of the ocean (i. e., the main thermocline and the deep sea). In these areas, the O_2 concentration is significantly depleted. The residual dissolved O_2 is enriched in ^{18}O because O_2 used for respiration was depleted in the heavy isotope [e. g., Kroopnick and Craig, 1976; Kroopnick, 1987]. We assume that 5% of gross production is remineralized in waters underlying the seasonal thermocline with an O_2 concentration equal to 40% of saturation and a $\delta^{18}\text{O}$ of O_2 enriched by 8 ‰ with respect to surface seawater. (The percent saturation

number is a rough estimate based on hydrographic sections of dissolved O_2 concentration. The $\delta^{18}O$ is estimated from the relationship between the $\delta^{18}O$ of dissolved O_2 in subsurface waters and the O_2 undersaturation [Bender, 1990].) The $\delta^{18}O$ of O_2 utilized by respiration in these waters will be equal to $\delta^{18}O$ (dissolved O_2) - ϵ_R , or 8-20 ‰ = 12 ‰. The effective respiratory isotope effect for seawater with respect to the $\delta^{18}O$ of ambient dissolved O_2 is then: $0.95 \times 19.3 \text{ ‰} + 0.05 \times 12 \text{ ‰} = 18.9 \text{ ‰}$.

Effect of Isotopic Exchange Between O_2 and CO_2 in the Stratosphere on the $\delta^{18}O$ of O_2

Recently, Gamo *et al.* [1989], Yung *et al.* [1991], Thiemens *et al.* [1991], and Wen and Thiemens [1993] demonstrated that isotopic exchange between O_2 and CO_2 occurs in the stratosphere and discussed its mechanism. $O(^1D)$, which ultimately derives from O_2 , reacts with CO_2 to form the CO_3 transition state complex. The symmetry properties of this complex are different for $C^{16}O_3$, $C^{16}O_2^{17}O$, and $C^{16}O_2^{18}O$. The difference in symmetry properties causes ^{17}O and ^{18}O to be enriched in the CO_2 that is produced by the decay of CO_3 . These heavy isotope enrichments are anomalous in the sense that they are the same for ^{17}O as for ^{18}O [Wen and Thiemens, 1993]. In almost all materials, the ^{17}O enrichments are roughly half those of ^{18}O , corresponding to the number of mass units by which ^{18}O and ^{17}O are heavier than the major isotope (^{16}O).

The magnitude of the stratospheric CO_2 enrichment, with respect to its tropospheric value, increases with height to about 11 ‰ between 27 and 35 km. The ^{17}O enrichment in stratospheric CO_2 is anomalous as it is nearly the same as the ^{18}O enrichment rather than half [Thiemens *et al.*, 1991], as noted above. Since O_2 is the ultimate source $O(^1D)$ reacting with CO_2 , the overall reaction represents isotopic exchange between CO_2 and O_2 in the stratosphere:



The effect of this transfer of ^{18}O from O_2 to CO_2 is to decrease the $\delta^{18}O$ of atmospheric O_2 . The question arises as to whether oxygen compounds other than O_2 can be the ultimate source of O exchanging with CO_2 . The answer is no, because the concentration of other O -bearing compounds (e. g., OH , H_2O) in the stratosphere is extremely low.

The quantitative effect of this exchange on the $\delta^{18}O$ of atmospheric O_2 can be estimated by scaling the stratospheric ^{18}O enrichment of CO_2 relative to the tropospheric value (taken as 11 ‰; [Thiemens *et al.*, 1991]) by the ratio of the CO_2 concentration of air to the O_2 concentration (353 ppmv / 210,000 ppmv), by the ratio of the turnover time of O_2 in air to the residence time of the troposphere with respect to exchange with the stratosphere (1.2 kyr / 56 years), and by a factor of -1 (since the effect on the $\delta^{18}O$ of O_2 is inverse to that for CO_2). (The residence time of O_2 in air is derived in Table 1, and the time assumed to be required for exchange of the stratosphere into the troposphere, 56 years, is from Schmidt and Khedim [1991].) One then computes that photochemical isotope exchange between CO_2 and O_2 in the stratosphere lowers the $\delta^{18}O$ of atmospheric O_2 by 0.4 ‰. There is obviously considerable uncertainty in this number.

M. Thiemens (personal communication, 1993) has pointed out that this estimated stratospheric $\delta^{18}O$ decrease of 0.4 ‰ should be accompanied by a $\delta^{17}O$ decrease of the same amount. These changes would lead to a non-mass-dependent fractionation of ^{17}O (defined as $\delta^{17}O - 0.5 \delta^{18}O$) of 0.2 ‰, well within the limits of detection. Data on the current value and its change through time would allow one to determine relative rates at which the non-mass-dependent anomaly is produced by photochemical reactions in the stratosphere and destroyed by photosynthesis and respiration.

Summary of the Discussion of the Dole Effect

On the basis of the preceding presentation (summarized in Table 1), we define three terms relevant to the discussion of the observed Dole effect and its change through time. The "terrestrial Dole effect" is defined as the hypothetical Dole effect when O_2 and CO_2 exchange isotopically only as the result of photosynthesis and respiration on land. The "marine Dole effect" is defined analogously. The "stratospheric diminution of the Dole effect" is the decrease in the $\delta^{18}O$ of O_2 in air due to isotope exchange between O_2 and CO_2 in the stratosphere. The various terms are summarized in Table 1. The estimate for marine production is calculated from seasonal net O_2 production as estimated by Keeling and Shertz [1992], assuming that gross O_2 production is four times seasonal net O_2 production. Other terms are taken from the literature or estimated as described above.

The calculated value for the Dole effect is then the productivity-weighted values of the terrestrial and marine Dole effects minus the stratospheric diminution: + 20.8 ‰. This value is considerably less than observed (+23.5 ‰). The difference between the expected value and the observed value reflects errors in our estimates and, conceivably, unrecognized processes. Recognizing the contribution of cyanide-resistant dark respiration would also increase the calculated value of the Dole effect.

We believe that the biggest current uncertainty lies in the $\delta^{18}O$ of leaf water. One can explain the observed value of the Dole effect by assuming a value of + 8.7 ‰, rather than + 4.4 ‰, for the $\delta^{18}O$ of leaf water, keeping everything else constant. Observations of the $\delta^{18}O$ of water in natural leaves from temperate and tropical forests during the growing season [Dongmann *et al.*, 1974; Forstel, 1978; Zundel *et al.*, 1978] are compatible with the higher value, but we should not attach too much importance to this observation because of natural variability in the $\delta^{18}O$ of leaf water. Nonetheless, we emphasize uncertainty in the $\delta^{18}O$ of leaf water for three reasons. First, the leaf water enrichment is the second most important cause (after respiratory isotope discrimination) of the stationary ^{18}O enrichment of O_2 relative to seawater. Second, the term is highly variable on all space and time scales and therefore very difficult to quantify. Third, and most important, there is no way to explain the Dole effect without invoking a terrestrial term (i. e., leaf water enrichment + ϵ_R) greater than the observed atmospheric O_2 $\delta^{18}O$ enrichment of + 23.5 ‰.

In the ensuing discussion, we regard the leaf water $\delta^{18}O$ estimate of Farquhar *et al.* [1993], 4.4 ‰, as a lower limit, since we doubt that uncertainties in other terms can account for the discrepancy from the calculated Dole effect of 20.8 ‰ and

the observed value. We regard the leaf water value of + 8.7 ‰, which brings the calculation into agreement with observation, as an upper limit, since part of the discrepancy between calculation and observation may be due to uncertainties indicated above (such as overestimation of marine production and neglect of the large ^{18}O fractionation associated with alternative pathway respiration).

Changes in the $\delta^{18}\text{O}$ of atmospheric O_2 and the Dole Effect During the Last 130 kyr

Records of the $\delta^{18}\text{O}$ of O_2 trapped in bubbles versus depth in the Vostok and GISP2 ice cores [Sowers *et al.*, 1991; Sowers *et al.*, 1993] allow us to reconstruct variations in the $\delta^{18}\text{O}$ of atmospheric O_2 and the Dole effect. For these reconstructions we need the following information: (1) data on the $\delta^{18}\text{O}$ of O_2 as a function of depth in the ice cores, (2) data on the $\delta^{15}\text{N}$ of N_2 for each ice core sample to correct the $\delta^{18}\text{O}$ of O_2 for gravitational fractionation [e. g., Craig *et al.*, 1988]; (3) estimates of the age of the gas in each ice core sample we studied, and (4) a record of the $\delta^{18}\text{O}$ of seawater ($\delta^{18}\text{O}_{\text{sw}}$) versus age during the relevant time period. We can then compute the Dole effect as a function of time by subtracting $\delta^{18}\text{O}_{\text{sw}}$ from $\delta^{18}\text{O}$ of contemporaneous atmospheric O_2 ($\delta^{18}\text{O}_{\text{atm}}$). We must calculate the Dole effect using compatible chronologies of deep sea sediments and ice cores. Relative errors in the chronostratigraphies introduce artifacts in the Dole effect curves since one would be subtracting $\delta^{18}\text{O}_{\text{atm}}$ at one time from $\delta^{18}\text{O}_{\text{sw}}$ at another. Chronological errors introduce errors into the calculated magnitude of changes in the Dole effect as well

as into the timing of the changes. The problem is obviously most acute at those times when $\delta^{18}\text{O}_{\text{sw}}$ and $\delta^{18}\text{O}_{\text{atm}}$ change rapidly.

For the seawater curve we use $\delta^{18}\text{O}$ data on the benthic foram *Uvigerina senticosa* from deep sea sediment core V 19-30 [Shackleton and Pisias, 1985]. The chronology is the SPECMAP chronology. We correct for the influence of bottom water temperature changes over the last 130 kyr on the foram $\delta^{18}\text{O}$ record as described by Sowers *et al.* [1993].

The chronology for the GISP2 ice core is based on annual layer counting throughout the interval relevant to this study [Alley *et al.*, 1993]. Ice age-gas age differences (Δage values) are calculated as described by Sowers *et al.* [1993b]. Δage is 220 ± 50 years for Holocene samples of GISP2 increasing to 630 ± 100 years for glacial samples. For the GISP2 samples, Δage adds only a very small uncertainty to the calculated value of the Dole effect.

The chronology we adopt for Vostok is that of Sowers *et al.* [1993]. They derived a timescale for the Vostok core by correlating their $\delta^{18}\text{O}_{\text{atm}}$ record with the $\delta^{18}\text{O}_{\text{sw}}$ curve (from the V 19-30 core) which had been placed into the SPECMAP timescale. This chronology appears to be robust between 0-20 kyr and 60-130 kyr. It is weak for the intervening period because the absence of large variations in $\delta^{18}\text{O}_{\text{atm}}$ and $\delta^{18}\text{O}_{\text{sw}}$ make correlation difficult. This chronology has several other weaknesses with respect to the objective of this paper. First, it is derived using the method of Martinson *et al.* [1982] with eight coefficients for the 133 kyr record. This limited number of coefficients leads to a "stiff" fit which does not reproduce the timing of high-frequency events very well. Second, we account for the lag in the response of $\delta^{18}\text{O}_{\text{atm}}$ to $\delta^{18}\text{O}_{\text{sw}}$ in a

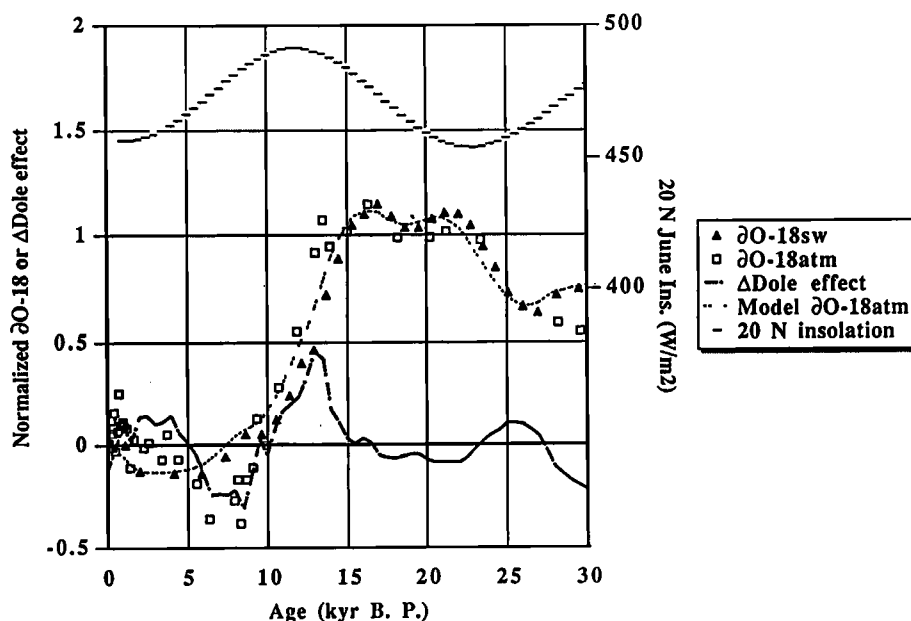


Figure 2. Experimental values of $\delta^{18}\text{O}_{\text{sw}}$ and $\delta^{18}\text{O}_{\text{atm}}$ from the GISP2 ice core versus age. The $\delta^{18}\text{O}_{\text{sw}}$ - age curve is from T. Sowers *et al.* (manuscript in preparation, 1994). The $\delta^{18}\text{O}_{\text{atm}}$ - age curve is from Sowers *et al.* [1993]. Also plotted are values of ΔDole effect (heavy line), calculated from data interpolated to 0.5-kyr intervals, and values $\delta^{18}\text{O}_{\text{atm}}$ predicted by a simple model assuming constant biogeochemical forcing (i. e., global carbon fluxes, leaf water fractionation, and photosynthetic and respiratory oxygen isotope fractionation as observed today).

way which, ideally, forces the Dole effect to be about + 0.2 ‰ relative to the present value during glacial terminations. Our calculated value of the Dole effect is larger during the terminations, possibly because $\delta^{18}\text{O}_{\text{atm}}$ appears to decrease more rapidly than $\delta^{18}\text{O}_{\text{sw}}$ during these events. Third, Sowers et al.'s approach to deriving a chronology for the Vostok core (by correlating $\delta^{18}\text{O}$ records) minimizes variations in the Dole effect throughout the record. Largely because of these problems, our curve of changes in the Dole effect as a function of time is provisional and will evolve as we get new data and improve the chronology of ice and sediment cores.

Values of $\delta^{18}\text{O}_{\text{atm}}$ for Vostok presented here are based on values of $\delta^{18}\text{O}$ of O_2 and $\delta^{15}\text{N}$ of N_2 reported by Sowers et al. [1993]. These results were from trapped gas samples from the 3G core which were extracted at the Laboratoire de Glaciologie et Geophysique de l'Environnement, in Grenoble, France, sealed in glass ampoules, and analyzed at the University of Rhode Island (URI). More recent results suggest an analytical problem with these results (which we call the "Grenoble Vostok 3G samples" in the ensuing discussion). First, $\delta^{18}\text{O}_{\text{atm}}$ values measured at URI for GISP2 samples (Figure 2) are systematically lower, by ~ 0.2 ‰, than values measured for Grenoble 3G samples. Second, $\delta^{18}\text{O}_{\text{atm}}$ values of samples from the Vostok 5G core, back to 30 kyr in age, extracted and measured at URI [Malaize, 1993], agree with values from the GISP2 core as opposed to Grenoble Vostok 3G data. An exhaustive review of our standard and methods calibrations revealed no errors except that the $\delta^{15}\text{N}$ of our standard may have shifted by 0.04 ‰ during the period the Grenoble Vostok 3G samples were analyzed. Such a shift would cause $\delta^{18}\text{O}_{\text{atm}}$ ($= \delta^{18}\text{O} - 2 \times \delta^{15}\text{N}$) to be 0.04 ‰ light for some samples and heavy by the same amount for others.

To address the question of calibration empirically, we prepared and analyzed nine samples from the 3G core at URI ("URI Vostok 3G samples"). In Figure 3, $\delta^{15}\text{N}$ and $\delta^{18}\text{O}$ are plotted versus depth, down to 700 m, for URI Vostok 3G samples, Grenoble Vostok 3G samples, and Vostok 5G samples (analyzed entirely at URI). $\delta^{18}\text{O}$ values for all three data sets are in excellent agreement throughout this interval. At depths above 650 m, where we have the greatest density of comparable samples, $\delta^{15}\text{N}$ values for the Vostok 5G samples

and URI Vostok 3G samples are systematically higher than Grenoble Vostok 3G samples by an average value of + 0.10 ‰. This offset most likely reflects an undetected artifact associated with sealing air samples in glass flasks, but we cannot confidently explain its origin.

In this paper we correct the $\delta^{15}\text{N}$ of all Grenoble Vostok 3G samples by + 0.10 ‰. This correction lowers $\delta^{18}\text{O}_{\text{atm}}$ of these samples by 0.20 ‰ and brings them into excellent agreement with $\delta^{18}\text{O}_{\text{atm}}$ values for GISP2.

Timing and Magnitude of Variations in the Dole Effect During the Last 130 KYR

GISP2 Record

In Figure 2 we plot $\delta^{18}\text{O}_{\text{sw}}$, $\delta^{18}\text{O}_{\text{atm}}$, and ΔDole effect versus age for the past 30 kyr, based on data from the GISP2 ice core. $\delta^{18}\text{O}_{\text{sw}}$ and $\delta^{18}\text{O}_{\text{atm}}$ are plotted with modern values normalized to zero. We define ΔDole effect as the difference between the Dole effect at time t in the past and the Dole effect today:

$$\Delta\text{Dole effect}(t) = \Delta\delta^{18}\text{O}_{\text{atm}}(t) - \Delta\delta^{18}\text{O}_{\text{sw}}(t). \quad (3)$$

We also calculate and plot the hypothetical value of $\delta^{18}\text{O}_{\text{atm}}$ if the biogeochemical forcing had been constant, that is to say, if the Dole effect had varied only because of the lag in the atmospheric response to a change in $\delta^{18}\text{O}_{\text{sw}}$. Finally, we plot average June insolation at 20°N . This term gives an index of summertime insolation changes in subtropical latitudes of the northern hemisphere. Relative variations of June insolation at 20°N are very close to those of summertime northern hemisphere insolation and inferred variations in low-latitude precipitation as deduced from modeling studies [Prell and Kutzbach, 1987]. We focus on low latitudes because many climate records from this region have a substantial amount of spectral density in the precessional frequencies (19 and 23 kyr) which correspond to the period of dominant variability in the Dole effect (see below).

Between about 16 and 21 kyr B. P. and again for the last 3 kyr, $\delta^{18}\text{O}_{\text{atm}}$ and $\delta^{18}\text{O}_{\text{sw}}$ are nearly constant. At these times of very different climate, the Dole effect is at steady state and

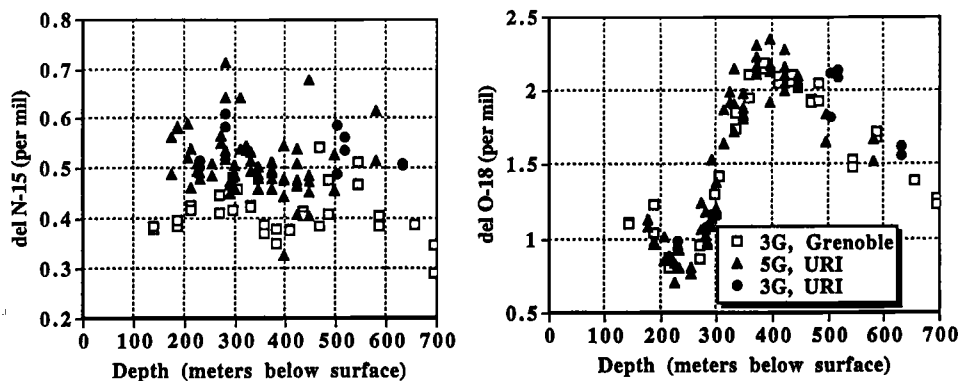


Figure 3. $\delta^{15}\text{N}$ of N_2 and $\delta^{18}\text{O}$ of O_2 versus depth in Vostok ice cores for samples from the 3G core extracted in Grenoble, the 3G core extracted at URI, and the 5G core extracted at URI. $\delta^{18}\text{O}$ values agree well but $\delta^{15}\text{N}$ values are offset by 0.1 ‰ between samples extracted in Grenoble and samples extracted at URI. For reference, nominal ice ages are 7.4 kyr at 200 m, 25 kyr at 400 m, and 42 kyr at 600 m depth.

very near to its present value. Δ Dole effect reaches a maximum at ~ 13 kyr B. P and a minimum at ~ 8 kyr B. P. Until about 11 kyr B. P., $\delta^{18}\text{O}_{\text{atm}}$ data are very close to the curve denoting constant biogeochemical fractionation. The Δ Dole effect maximum at 13 kyr BP may largely reflect the 1.2 kyr turnover time of atmospheric O_2 and the associated lag in the response of $\delta^{18}\text{O}_{\text{atm}}$ to changes in $\delta^{18}\text{O}_{\text{sw}}$. On the other hand, the minimum at about 8 kyr B. P. is undoubtedly due to changing isotope fractionations associated with transient processes of the biosphere and hydrosphere. This minimum in Δ Dole effect may be related to the inferred maximum in low-latitude precipitation at ~ 10 kyr B. P. and the intensification of the monsoon at this time [Prell and Kutzbach, 1987], reforestation of the temperate northern hemisphere and the associated minimum in atmospheric pCO_2 [Neftel et al., 1988], and/or to transients in oceanic circulation and productivity at the end of the last ice age.

Vostok Record

Δ Dole effect calculated for Vostok is plotted versus age, along with average June insolation at 20°N , in Figure 4a.

Δ Dole effect varies with the 23-kyr precession cycle (Table 2). This periodicity is reflected in maxima in the Dole effect at about 128 - 133 kyr, 108 kyr, 84 - 91 kyr, 58 kyr, 35 kyr, 25 kyr, and 13 kyr B. P. The 25-kyr maximum in Δ Dole effect is anomalous, because it does not correspond to a maximum in insolation. However, it would disappear if the age assigned to the $\delta^{18}\text{O}_{\text{atm}}$ minimum at 30 kyr were decreased by about 4 kyr to align the minimum with the $\delta^{18}\text{O}_{\text{sw}}$ minimum at 26 kyr [Sowers et al., 1993]. Such a change would also bring the chronology of Sowers et al. into better agreement with other chronologies proposed for the Vostok ice core during this time period. At the present time, we merely note that sample resolution between 25 and 47 kyr in the Vostok ice core is not good enough to generate a reliable chronology based on $\delta^{18}\text{O}_{\text{atm}}$ alone.

Variations of Δ Dole effect, as calculated from the Sowers et al. [1993] chronology, are highly coherent with, and nearly in phase with, insolation. The phasing is similar to that observed in the GISP2 record. Maxima in Δ Dole effect occur at times of maximum insolation. Minima in Δ Dole effect lead minima in insolation (at 100 kyr, 77 kyr, and 7 kyr) or occur in phase (115 kyr, 46 kyr). At the time of termination II, variations in Δ Dole effect may largely reflect the 1.2-kyr turnover time of O_2 in air, as for GISP2 (see above). However, for other events the coincidence of the Δ Dole effect maxima with respect to insolation cannot be due to the lag in the atmospheric response. The reason for this is that the change in $\delta^{18}\text{O}_{\text{sw}}$ is too small, and the calculated atmospheric response is too rapid, to induce a change of the magnitude observed. At times other than the terminations, variability in Δ Dole effect must either be biogeochemical or an artifact of the chronology.

Unfortunately, errors in the chronology of Vostok with respect to SPECMAP can induce very large changes in apparent values of Δ Dole effect. We illustrate this with Figures 4b and 4c. In Figure 4b we plot Δ Dole effect after holding the chronology of SPECMAP constant but decreasing all Vostok ages by 3 kyr. This change in the chronology of Vostok reduces noise and greatly increases the coherency

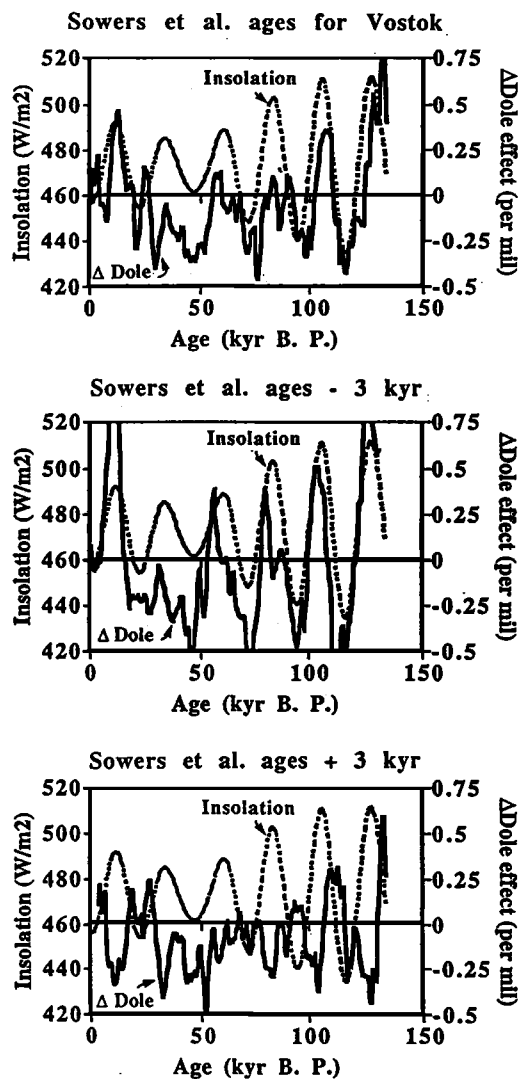


Figure 4. Δ Dole effect and integrated June insolation versus SPECMAP age at 20°N . (a) Sowers et al. [1993] chronology for $\delta^{18}\text{O}_{\text{atm}}$ (Vostok) versus age, (b) 3 kyr subtracted from Sowers et al. age, (c) 3 kyr added to Sowers et al. age. The SPECMAP chronology is used throughout for the $\delta^{18}\text{O}_{\text{sw}}$ curve. This figure illustrates the high coherency between Δ Dole effect and insolation for the Sowers et al. chronology but also shows that the calculated values of Δ Dole effect are rather sensitive to relative errors in the ages of the Vostok and SPECMAP records.

between Δ Dole effect and insolation (Table 2b). In Figure 4c we calculate Δ Dole effect after holding the chronology of SPECMAP constant but increasing all Vostok ages by 3 kyr. This change in the Vostok chronology adds a great deal of noise to the Δ Dole record, reduces the coherency between Δ Dole and insolation (Table 2), and changes the phase by about 130° (Table 2). The coherent nature of the Δ Dole effect and insolation curves is robust, suggesting that changes in the Dole effect are related to insolation variations. However, we do not yet understand the chronology of Vostok well enough to determine phase differences between Δ Dole effect and other climate variables.

Table 2a. Spectral Analysis

Variable	Period, kyr	f statistic
Δ Dole effect	22.8	0.93
	13.2	0.84
CH ₄	13.6	0.85
	22.3	0.99
	14.5	0.93
$\delta^{18}\text{O}_{\text{atm}}$	41.8	0.92
	21.6	0.93
	18.1	0.81

Analyses were run using the multi-taper method. Results are given only for frequencies where the f-statistic was greater than 0.8.

Table 2b. Cross Spectral Analysis

Variables	Period, kyr	Coherency	Phase, kyr
CH ₄ - Δ Dole effect	21.5	0.82	5.1 (CH ₄ leads)
Δ Dole-insolation	22.1	0.92	2 (insolation leads)
Δ Dole-insolation; subtract 3 kyr from SPECMAP gas age	20.6	0.96	1 (Δ Dole leads)
Δ -insolation; add 3 kyr to gas age	19.4	0.79	7 (insolation leads)

Cross spectral analyses run using the Blackman-Tuckey method.

Results are given only when the coherency was greater than 0.8.

"Insolation" refers to average June insolation at 20° N.

Why Does Δ Dole Effect Vary With the Periodicity of Precession?

Most variability in insolation at low latitudes is associated with precession, while at high latitudes tilt also becomes important. The strong precession signal in Δ Dole effect suggests that variations in the Dole effect are probably linked to low latitude processes. There are strong precession signals in both marine and continental proxy records from low latitudes (e. g., *Clemens and Prell*, 1990; *Chappellaz et al.*, 1990). In principle, we could use the phasing of proxy signals and Δ Dole effect to test whether certain climate variables influence Δ Dole effect. For example, if low-latitude ocean productivity changes were driving changes in the Dole effect, we would expect these two records to be about 180° out of phase (high oceanic productivity leads to lower Dole effect). In practice, we do not know phasing well enough to use this criterion.

Most marine proxy records show variability in the precession band [e. g., *Imbrie et al.*, 1992], but those showing dominant variability in the precession band are rare during the late Pleistocene due to the predominant influence of the ice sheets on glacial - interglacial climate cycles with their associated obliquity signal. They are restricted to the equatorial Atlantic [*McIntyre et al.*, 1989; *Molfin and McIntyre*, 1990] and the equatorial Indian Oceans [*Clemens and Prell*, 1990]. It seems possible but unlikely that variability of primary productivity in these restricted areas of the ocean would have dominated changes in the Dole effect. We therefore look to changes in the continental biosphere to explain variations in the Dole effect.

The CH₄ concentration is the simplest indicator of large-scale variations in the continental biosphere. CH₄ variations during the last 150 kyr B. P., measured in the Vostok ice core by *Chappellaz et al.* [1990], vary by a factor of 2, with most variability concentrated in the precession band. *Chappellaz et al.* [1990] linked CH₄ maxima to maxima in paleoprecipitation as estimated by *Prell and Kutzbach* [1987]. The connection results from the fact that precipitation is linked to evaporation at low latitudes and therefore to insolation. Increased precipitation stimulates the activity of the biosphere, particularly in wetlands which are the major CH₄ source.

Increased precipitation would also cause the Dole effect to change, because continental productivity would increase and the leaf water enrichment would vary. The effect of increased terrestrial production would be to raise the Dole effect by increasing the production rate of ¹⁸O - enriched photosynthetic O₂ coming from the land biosphere. The effect of increased wetness and humidity is to lower the Dole effect by diminishing the leaf water enrichment. The phasing between insolation and the Dole effect is unclear. Improved chronologies for deep-sea sediments, together with modeling studies of terrestrial productivity and leaf water enrichments such as those of *Farquhar et al.* [1993], will allow us to improve our record of past changes in the Dole effect and our understanding of their biogeochemical implications.

Why is Variability in the Dole Effect so Small?

As shown above, Δ Dole effect varies with a period of ~23 kyr, suggesting a response by the low-latitude biosphere linked to insolation changes induced by precession. This periodicity notwithstanding, Δ Dole effect has been remarkably constant. The mean and standard deviation, based on Vostok, has been - 0.05 ‰ ± 0.24 ‰, as calculated from our record interpolated to 0.5 kyr intervals.

Even more striking is the similarity of the Dole effect between glacial and interglacial times. The Δ Dole effect during the last glacial maximum was close to 0 ‰ (Figures 2 and 4a). The change in Δ Dole effect between the penultimate glacial maximum (stage 6) and the last interglacial (stage 5e) was also very small: $\delta^{18}\text{O}_{\text{sw}}$ decreased by ~1.2 ‰ during the termination, and $\delta^{18}\text{O}_{\text{atm}}$ decreased by ~1.4 ‰ (Figure 4a) [*Sowers et al.*, 1991]. There may indeed have been a large change in Δ Dole effect during termination II similar to the change we observed during termination I (Figure 4a). However, between full glacial and full interglacial climates the

change in the Dole effect was very small, despite the drastic changes in the water cycle and biogeography. Why has the Dole effect been so constant?

Answering this question is impossible at the present time, but we can probably recognize the key influences. The influence of the land biosphere during glacial times would have been different from that today in the following ways:

1. During glacial times, climate was colder and drier and $p\text{CO}_2$ was lower. Much of the temperate and polar regions were covered by ice. Cold climate, low precipitation, and extensive ice cover would all lower glacial productivity; Meyer [1988] has estimated that terrestrial net primary production during the last glacial maximum (LGM) was ~75 % of the present value. If everything else were constant, this change would have lowered the rate of photosynthesis by the land biosphere and hence the Dole effect.

2. Drier climates probably would have caused more extensive leaf water fractionation, which would cause the Dole effect to rise.

3. Meyer's estimates of production during the LGM do not include the influence of lower atmospheric CO_2 on glacial productivity. If production varies by $\text{CO}_2^{0.3}$ [Keeling *et al.*, 1989], glacial productivity on land would have decreased by an additional factor of $(180 \text{ ppmV}/280 \text{ ppmV})^{0.3}$, or 0.88, further lowering the Dole effect. The true effect of lower CO_2 would be much less if nutrients ultimately limit production.

4. The lower CO_2 concentration in air during the last glacial period would have caused the ratio of photorespiration/carboxylation (ϕ) to increase from a calculated preindustrial value of 0.31-0.48 during the LGM, according to (2). On the other hand, the colder glacial

temperatures would have caused ϕ to decrease by a factor of 0.88. This value is calculated taking Q10 for the ratio of oxidation to carboxylation to be 0.57 [Collatz *et al.*, 1991]. Overall, we expect ϕ to have risen from a calculated preindustrial value of 0.31-0.42 during the LGM. Other things being equal, this increase in ϕ would have caused the Dole effect to rise for two reasons. First, it heralds an increase in terrestrial gross production, which raises the Dole effect because of the leaf water enrichment. Second, there is a relatively large isotope effect associated with photorespiration.

On the marine side, isotope discrimination probably was similar during the LGM and today. Changes in the oceanic contribution to the Dole effect would have come from changes in marine production. Oceanic production during the LGM differed from that of today in two ways:

1. A number of studies have examined the fertility of the deep oceans during the last glacial maximum. The most ambitious of these are the studies using sedimentary organic carbon accumulation rates [e. g., Sarnthein *et al.*, 1988] and foraminiferal transfer functions [e. g., Mix, 1989a, b]. The former approach has given contradictory results in different oceanic areas (hardly surprising, since there is no reason to suspect that productivity would change uniformly throughout the ocean), but has generally indicated higher glacial values [Sarnthein *et al.*, 1988]. The same is true of the transfer function approach in the Atlantic [Mix, 1989a]. On average, both approaches indicate higher deep ocean productivity during the last glacial maximum. However, we note that neither approach is free of problems, and results from these two approaches are not mutually consistent, sometimes

Table 3. Literature Estimates of Paleoproductivity During the Last Glacial Maximum

Citation	Study region	Paleoproductivity
<i>Sarnthein et al.</i> [1988]	Global	~1.25
<i>Marino et al.</i> [1992]	Global	Lower in polar ocean, higher elsewhere
Equatorial Pacific		
<i>Thunnell et al.</i> [1992]	South China Sea	2
<i>Archer</i> [1991]	Central equatorial Pacific	~2
<i>Lyle</i> [1988]	Equatorial Pacific	>1 (~2?)
<i>Pedersen et al.</i> [1991]	Eastern equatorial Pacific	~3
<i>Lyle et al.</i> [1992]	Central equatorial Pacific	Glacial upwelling more intense
Oregon Coast		
<i>Dymond et al.</i> [1992]		0.5
Atlantic Ocean		
<i>Mix</i> [1989a]	Entire Atlantic Ocean	~1.25
<i>Molfino and McIntyre</i> [1990]	Equatorial Atlantic	Glacial upwelling more intense
<i>Lyle</i> [1988]	Equatorial Atlantic	>1 (~2?)
Southern Ocean		
<i>Charles et al.</i> [1991]	Atlantic and Indian regions	~1 but meridionally variable

giving downcore variations of production that are very different in magnitude [e. g., *Sarnthein et al.*, 1992, Figures 3, 4, and 6], which are uncorrelated within individual cores [*Sarnthein et al.*, 1992, Figure 7], or which are anticorrelated over long depth intervals [Mix, 1989b]. A wide range of other approaches has been used to estimate paleoproductivity during the last glacial maximum. Results, summarized in Table 3, are unanimous in concluding that equatorial Pacific productivity was higher during the last glacial maximum. Elsewhere, the sense of productivity variations differed from one region to another. However, the general sense is that open ocean production was higher during the last glacial maximum. This would cause a drop in the Dole effect.

2. The continental shelves, which are highly productive, account for ~7 % of the area of the sea floor and ~20 % of total oceanic production [Walsh, 1991]. During glacial times, most of the shelf area would have been exposed and would thus not have contributed to oceanic production. In addition, production on the shallow part of the continental slope would have been suppressed by extensive sea ice cover in the Arctic and Antarctic and by isostatic depression of the landward portion of the continental slope bordering glaciated areas. These processes would have caused global ocean productivity to fall and the Dole effect to rise.

In summary, we can conclude that the Dole effect has been rather constant during the last 130 kyr because environmental changes have induced roughly offsetting changes in oxygen isotope fractionation of O₂. We can recognize the nature of these changes, but we cannot account quantitatively for their varying influence on the Dole effect.

Summary and Conclusions

Our ability to explain the magnitude of the contemporary Dole effect is a measure of our understanding of the global cycles of oxygen and water. A variety of recent studies have improved our understanding of many of the principles governing oxygen isotope fractionation during photosynthesis and respiration, of production rates on land and in the oceans, and of the effect of stratospheric photochemistry on the δ¹⁸O of O₂ in air. However, our attempt to quantitatively account for the Dole effect in terms of these principles was not very successful: We estimated the present-day Dole effect to be + 20.8 ‰ compared to the observed value of + 23.5 ‰. The agreement is considerably worse than it might appear given the fact that respiratory isotope fractionation alone must account for ~20 ‰ of the stationary enrichment of the ¹⁸O of O₂ compared with seawater. Our underestimate of the true magnitude most likely reflects an underestimate in the leaf water enrichment, failure to consider respiration by the alternative (cyanide resistant) pathway, and perhaps an overestimate of marine production.

We have used data on the δ¹⁸O of O₂ from the Vostok ice core to calculate changes in the Dole effect during the last 130 kyr. Our results show that variations in the Dole effect have been relatively small during most of the last glacial - interglacial cycle. These small changes are not consistent with large glacial increases in global oceanic productivity. Changes in the Dole effect vary with the period of precession, and we conclude that they largely reflect changes in terrestrial

productivity and leaf water fractionation associated with variations in humidity and tropical precipitation.

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References

- Alley, R. B., D. Meese, C. A. Shuman, A. J. Gow, K. Taylor, M. Ram, E. Waddington, J. W. C. White and P. Mayewski, Abrupt accumulation increase at the Younger Dryas termination in the GISP2 ice core, *Nature*, 362, 527-529, 1993.
- Archer, D., Equatorial Pacific calcite preservation cycles: Production or dissolution?, *Paleoceanography*, 6, 561-571, 1991.
- Bender, M., The δ¹⁸O of dissolved O₂ in seawater: S unique tracer of circulation and respiration in the deep sea, *J. Geophys. Res.*, 95, 22,243-22,252, 1990.
- Bender, M., L. Labeyrie, D. Raynaud, and C. Lorius, Isotopic composition of atmospheric O₂ in ice linked with deglaciation and global primary productivity, *Nature*, 318, 349-352, 1985.
- Benson, B., and D. Krause, Jr., The concentration and isotopic fractionation of oxygen dissolved in fresh water and seawater in equilibrium with the atmosphere, *Limnol. Oceanogr.*, 29, 620-632, 1984.
- Berry, J., Biosphere, atmosphere, ocean interactions: A plant physiologist's perspective, in *Primary Productivity and Biogeochemical Cycles in the Sea*, edited by P. G. Falkowski and A. D. Woodhead, pp. 441-454, Plenum, New York, 1992.
- Boyle, E., The role of vertical chemical fractionation in controlling Late Quaternary atmospheric carbon dioxide, *J. Geophys. Res.*, 93, 15,701-15,715, 1988.
- Broecker, W. S., The cause of the glacial to interglacial atmospheric CO₂ change: A polar alkalinity hypothesis, *Global Biogeochem. Cycles*, 3, 215-239, 1989.
- Chappellaz, J., J. M. Barnola, D. Raynaud, Y. S. Korotkevich, and C. Lorius, Ice core record of atmospheric methane over the past 160,000 years, *Nature*, 345, 127-131, 1990.
- Charles, C. D., P. N. Froelich, M. A. Zibello, R. A. Mortlock, and J. J. Morley, Biogenic opal in southern ocean sediments over the last 450,000 years: Implications for surface water chemistry and circulation, *Paleoceanography*, 6, 697-728, 1991.
- Clemens, S., and W. Prell, Late Pleistocene variability of Arabian sea summer monsoon winds and continental aridity: Eolian records from the lithogenic component of deep-sea sediments, *Paleoceanography*, 5, 109-145, 1990.
- COHMAP members, Climatic changes of the last 18,000 years: Observations and model simulations, *Science*, 241, 1043-1052, 1988.
- Collatz, G. T., J. T. Ball, C. Grivet, and J. A. Berry, Physiological and environmental regulation of stomatal conductance, photosynthesis and transpiration: A model that includes a laminar boundary layer, *Agric. and For. Meteorol.*, 54, 107-136, 1991.
- Craig, H. and L. Gordon, Deuterium and oxygen-18 variation in the ocean and marine atmosphere. paper presented at Conference on Stable Isotopes in Oceanography, International Atomic Energy Agency, Pisa, 1965.
- Craig, H., Y. Horibe, and T. A. Sowers, Gravitational separation of gases and isotopes in polar ice caps, *Science*, 242, 1675-1678, 1988.
- Dole, M., G. Lane, D. Rudd, and D. A. Zaukelies, Isotopic composition of atmospheric oxygen and nitrogen, *Geochim. Cosmochim. Acta*, 6, 65-78, 1954.
- Dongmann, G., The contribution of land photosynthesis to the stationary

- enrichment of ^{18}O in the atmosphere, *Radiat. Environ. Biophys.*, **11**, 219-225, 1974.
- Dongmann, G., H. W. Nurnberg, H. Forstel and K. Wagener, On the Enrichment of H_2^{18}O in the leaves of transpiring plants, *Radiat. Environ. Biophys.*, **11**, 41-52, 1974.
- Dymond, J., E. Suess, and M. Lyle, Barium in deep-sea sediment: A geochemical proxy for paleoproductivity, *Paleoceanography*, **7**, 163-181, 1992.
- Epstein, S., and L. Zeiri, Oxygen and carbon isotopic composition of gases respired by humans, *Proc. Natl. Acad. Sci. U. S. A.*, **85**, 1727-1731, 1988.
- Farquhar, G. D., S. von Caemmerer, and J. A. Berry, A biochemical model of photosynthetic CO_2 assimilation in leaves of C_3 species, *Planta*, **149**, 78-90, 1980.
- Farquhar, G. D., J. Lloyd, J. A. Taylor, L. Flanagan, J. P. Syvertsen, K. T. Hubick, S. Wong, and J. R. Ehleringer, Vegetation effects on the isotope composition of oxygen in atmospheric CO_2 , *Nature*, **363**, 439-443, 1993.
- Farris, F. and B. R. Strain, The effects of water-stress on leaf H_2^{18}O enrichment, *Rad. and Environm. Biophys.*, **15**, 167-202, 1978.
- Forstel, H., The enrichment of ^{18}O in leaf water under natural conditions, *Radiat. Environ. Biophys.*, **15**, 323-344, 1978.
- Gamo, T., M. Tsutsumi, H. Sakai, T. Nakazawa, M. Tanaka, H. Honda, H. Kubo, and T. Itoh, Carbon and oxygen isotopic ratios of carbon dioxide of a stratospheric profile over Japan, *Tellus*, **41B**, 127-133, 1989.
- Guy, R. D., J. A. Berry, M. L. Fogel and T. C. Hoering, Differential fractionation of oxygen isotopes by cyanide-resistant and cyanide-sensitive respiration in plants, *Planta*, **177**, 483-491, 1989.
- Guy, R. D., J. A. Berry, M. L. Fogel, D. H. Turpin and H. G. Weger, Fractionation of the stable isotopes of oxygen during respiration by plants-The basis of a new technique to estimate partitioning to the alternative path, in *Biochemical and Physiological Aspects of Plant Respiration*, edited by H. Lambers and L. H. W. van der Plas, pp. 443-453, SPB Academic Publishing, The Hague, Netherlands, 1992.
- Guy, R. D., M. L. Fogel, and J. A. Berry, Photosynthetic fractionation of stable isotopes, *Plant Physiol*, **101**, 37-47, 1993.
- Horibe, Y., K. Shigehara, and C. C. Langway Jr., Chemical and isotopic composition of air in a Greenland ice core, *Earth Planet. Sci. Lett.*, **73**, 207-210, 1985.
- Imbrie, J., et al., On the structure and origin of major glaciation cycles, 1, Linear responses to Milankovitch Forcing, *Paleoceanography*, **7**, 701-738, 1992.
- Keeling, C. D., R. B. Bacastow, A. R. Carter, S. C. Piper, T. P. Whorf, M. Heimann, W. Mook, and H. Roeloffzen, A three-dimensional model of atmospheric CO_2 transport based on observed winds: 1. Analysis of observational data, in *Aspects of Climate Variability in the Pacific and the Western Americas*, Geophys. Monogr. Ser. Vol. 55, edited by D. H. Peterson, pp. 165-236, American Geophysical Union, Washington, D. C., 1989.
- Keeling, R., and S. Shertz, Seasonal and interannual variations in atmospheric oxygen and implications for the global carbon cycle, *Nature*, **358**, 723-727, 1992.
- Kiddon, J., M. Bender, J. Orchardo, J. Goldman, D. Caron, and M. Dennett, Isotopic fractionation of oxygen by respiring marine organisms, *Global Biogeochem. Cycles*, **7**, 679-694, 1993.
- Knox, F., and M. McElroy, Changes in atmospheric CO_2 ; influence of marine biota at high latitudes, *J. Geophys. Res.*, **89**, 4629-2637, 1984.
- Kroopnick, P. M., Oxygen 18 in dissolved oxygen, in *Geosecs Atlantic, Pacific, and Indian Ocean Expeditions: Shorebased Data and Graphics*, Vol. 7, edited by H. G. Ostlund, H. Craig, W. S. Broecker, and D. Spencer, pp. 3 and 27-182, U. S. Government Printing Office, Washington, D. C., 1987.
- Kroopnick, P., and H. Craig, Atmospheric oxygen: Isotopic composition and solubility fractionation, *Science*, **175**, 54-55, 1972.
- Kroopnick, P., and H. Craig, Oxygen isotope fractionation in dissolved oxygen in the deep sea, *Earth Planet. Sci. Lett.*, **32**, 375-388, 1976.
- Lane, G., and M. Dole, Fractionation of oxygen isotopes during respiration, *Science*, **123**, 574-576, 1956.
- Lyle, M., Climatically forced organic carbon burial in equatorial Atlantic and Pacific Oceans, *Nature*, **335**, 529-532, 1988.
- Lyle, M. W., F. G. Prah, and M. A. Sparrow, Upwelling and productivity changes inferred from a temperature record in the central equatorial Pacific, *Nature*, **355**, 812-815, 1992.
- Malaize, Bruno, thesis.
- Marino, B., M. McElroy, R. Salawitch, and W. Spaulding, Glacial-to-interglacial variations in the carbon isotopic composition of atmospheric CO_2 , *Nature*, **357**, 461-466, 1992.
- Martinson, D. G., W. Menke, and P. Stoffa, An inverse approach to signal correlation, *J. Geophys. Res.*, **87**, 4807-4818, 1982.
- McIntyre, A., W. Ruddiman, K. Karlin, and A. Mix, Surface water response of the equatorial Atlantic Ocean to orbital forcing, *Paleoceanography*, **4**, 19-55, 1989.
- Meyer, M. K., Net primary productivity estimates for the last 18,000 years evaluated from simulations by a global climate model, M. S. dissertation, Univ. of Wis. Madison, 1988.
- Mix, A., Productivity of the Pleistocene Atlantic ocean estimated from foraminiferal species: Implications for paleo- pCO_2 , *Nature*, **337**, 541-544, 1989a.
- Mix, A., Pleistocene paleoproductivity: Evidence from organic carbon and foraminiferal species, in *Productivity of the Ocean: Present and Past*, edited by W. H. Berger, V. Smetacek, and G. Wefer, John Wiley, New York, 1989b.
- Molfin, B. and A. McIntyre, Precessional forcing of nutricline dynamics in the equatorial Atlantic, *Science*, **249**, 766-769, 1990.
- Nefel, A., H. Oeschger, T. Staffelbach, and B. Stauffer, CO_2 record in the Byrd ice core 50,000-5,000 years BP, *Nature*, **331**, 609-611, 1988.
- Pedersen, T. F., B. Nielsen, and M. Pickering, Timing of late Quaternary productivity pulses in the Panama Basin and implications for atmospheric CO_2 , *Paleoceanography*, **6**, 657-677, 1991.
- Prell, W. L. and J. E. Kutzbach, Monsoon variability over the past 150,000 years, *J. Geophys. Res.*, **92**, 8411-8425, 1987.
- Sarnthein, M., K. Winn, J.-C. Duplessy, and M. R. Fontugne, Global variations of surface ocean productivity in low and mid latitudes: influence on CO_2 reservoirs of the deep ocean and atmosphere during the last 21,000 years, *Paleoceanography*, **3**, 361-399, 1988.
- Sarnthein, M., U. Pflaumann, R. Ross, R. Tiedemann, and K. Winn, Transfer functions to reconstruct ocean paleoproductivity: a comparison, *Geol. Soc. London*, **63**, 411-427, 1992.
- Schmidt, U., and A. Khedim, In situ measurements of carbon dioxide in the winter Arctic vortex and at midlatitudes: An indicator of the 'age' of stratospheric air, *Geophys. Res. Lett.*, **18**, 763-766, 1991.
- Shackleton, N. J., and N. G. Pisias, Atmospheric carbon dioxide, orbital forcing, and climate, in *The Carbon Cycle and Atmospheric CO_2 : Natural Variations Archean to Present*, *Geophys. Monogr. Ser.*, Vol. 32, edited by E. T. Sundquist and W. S. Broecker, pp. 303-317, AGU, Washington, D. C., 1985.
- Sowers, T., M. Bender, D. Raynaud, Y. S. Korotkevich, and J. Orchardo, The $\delta^{18}\text{O}$ of atmospheric O_2 from air inclusions in the Vostok ice core: Timing of CO_2 and ice volume changes during the penultimate deglaciation, *Paleoceanography*, **6**, 679-696, 1991.
- Sowers, T., M. Bender, L. Labeyrie, D. Martinson, J. Jouzel, D. Raynaud, and Y. S. Korotkevich, 135,000 year Vostok - SPECMAP common temporal framework, *Paleoceanography*, **8**, 737-766, 1993.
- Stevens, C., R. Schultz, C. Van Baalen, and P. Parker, Oxygen isotope fractionation during photosynthesis in a blue-green and a green alga, *Plant Physiol.*, **56**, 126-129, 1975.
- Thiemens, M., T. Jackson, K. Mauersberger, B. Schueler, and J. Morton, Oxygen isotope fractionation in stratospheric CO_2 , *Geophys. Res. Lett.*, **18**, 669-672, 1991.
- Thunell, R. C., M. Qingmin, S. Calvert and T. F. Pedersen, Glacial-Holocene biogenic sedimentation patterns in the south China Sea:

- Productivity variations and surface water $p\text{CO}_2$, *Paleoceanography*, 7, 143-162, 1992.
- Walsh, J. J., Importance of continental margins in the marine biogeochemical cycling of carbon and nitrogen, *Nature*, 350, 53-55, 1991.
- Wen, J., and M. H. Thiemens, First multi-isotope study of the $\text{O}(^1\text{D}) + \text{CO}_2$ exchange and stratospheric consequences. *J. Geophys. Res.*, 98, 12,801-12,808, 1993.
- Yakir, D., M. DeNiro, and P. Rundel, Isotopic inhomogeneity of leaf water: Evidence and implications for the use of isotopic signals transduced by plants, *Geochim. Cosmochim. Acta*, 53, 2769-2773, 1989.
- Yung, Y., W. De Mone, and J. P. Pinto, Isotopic exchange between carbon dioxide and ozone via $\text{O}(^1\text{D})$ in the stratosphere, *Geophys. Res. Lett.*, 18, 13-16, 1991.
- Zundel, G., W. Miekeley, B. M. Grisi, and H. Forstel, The H_2^{18}O enrichment in the leaf water of tropic trees: comparison of species from the tropical rain forest and the semi-arid region in Brazil, *Radiat. Environ. Biophys.*, 15, 203-212, 1978.

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