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# Magnetically driven superhydrophobic/superoleophilic graphene-based polyurethane sponge for highly efficient oil/water separation and demulsification

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**Magnetically driven superhydrophobic/superoleophilic graphene-based** 1  
**polyurethane sponge for highly efficient oil/water separation and** 2  
**demulsification** 3

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**Abstract** 24

Cost-effective and efficient sorbents with emulsion-breaking capability, especially surfactant-stabilized emulsions, are required to remediate oil spills from water. Herein, *in situ* growth of magnetic  $Mn_{0.01}Co_{0.90}Fe_3O_4$  (MCFO) nanoparticles (NPs) on reduced graphene oxide (RGO) was performed to prepare MCFO/RGO nanocomposite *via* a simple solvothermal method. The nanocomposite was coated onto a polyurethane (PU) sponge to fabricate a novel magnetic, superhydrophobic/superoleophilic and flame-retardant material. The prepared magnetic MCFO/RGO/PU was superhydrophobic with a contact angle of  $165 \pm 1^\circ$ . Compared with neat PU, the MCFO/RGO/PU sponge could be effectively driven by a magnet for oil collection in water. More importantly, the MCFO/RGO/PU sponge could effectively collect oil droplets from surfactant-free and stabilized toluene/water emulsions. In addition, the fabricated sponge exhibited excellent performance for oils and organic solvents sorption, achieving up to 131 times of its own weight and retaining its ability even after 10 sorption cycles. Interestingly, the thermal stability and fire retardancy of the prepared PU sponge were improved with respect to neat PU.

**Keywords:**  $Mn_{0.01}Co_{0.90}Fe_3O_4$ /reduced graphene oxide; magnetic polyurethane sponge; superhydrophobic/superoleophilic; demulsification capacity; flame retardant

## 1. Introduction

48

Nowadays, the production of oily wastewater discharged from domestic and industrial activities is extremely increasing. Therefore, the separation of oil/water mixtures and treating oily wastewater resources have become an urgent world challenge. Generally, oil/water mixtures are classified into two categories including immiscible oil-water and miscible oil-water mixtures (oil-in-water (O/W) and water-in-oil (W/O) emulsions). Due to the small size ( $d < 20 \mu m$ ) and high stability of emulsion droplets, their separation is more challenging and difficult than immiscible oil/water mixtures with average droplet size up to  $150 \mu m$  [1]. As a result, much efforts have been devoted to design novel, low cost, reusable, and recyclable sorbents for oil/water separation.

The most common methods used for the separation of oil/water mixtures are sorption and filtration [2]. Nanocomposites, gels, powders, particles, and porous media are usually used as sorbent materials, while the common filtration materials applied in oil-water separation are films, filters, membranes, and meshes. Although filtration has been used as a successful method in industry, contaminants could not be effectively collected from mixtures. Sorbents, especially 3-dimensional (3D) porous materials, can easily adsorb contaminants from wastewater resources [3]. They are not expensive and there is no need of any specialized equipment to clean up the contaminants. Both natural including cotton fibers, kapok fibers, wool and synthetic sponges (e.g. polyurethane (PU) and melamine sponges), polystyrene fibers, and carbon nanotubes (CNTs) have been widely used as sorbents [2]. Although, natural materials offer some excellent advantages such as high abundance and low cost, they suffer from some shortcomings like flammability, non-reusability, introduction of new pollutants into the contaminated system, and poor sorption capacity. In addition, they are not suitable for separation of emulsified oil/water mixtures, especially in surfactant-stabilized emulsions [4]. In contrast, 3D porous sorbents like PU sponge with its specific features

represents an appealing alternative. PU sponge could act as a promising candidate for oil-water separation, because of its outstanding properties like large surface area, 3D skeleton, porous structure, remarkable flexibility, good sorption capacity, promising reusability, and inter-connected matrix [5-9]. Additionally, for effective oil/water separation, superhydrophobicity (surface with a water contact angle (WCA) higher than  $150^\circ$ ) is crucial. Since PU is slightly hydrophobic, its non-wettability should be enhanced [10]. In addition, PU sponge is easily ignited and generates toxic gases [11]. These drawbacks hamper significantly the practical applications of PU sponge in oil-water separation [12]. Therefore, it is necessary to alter surface wettability of PU and make it suitable for oil/water separation processes such as emulsions (W/O and O/W) and highly viscous oil/water mixtures. There are several reports for the preparation of superhydrophobic/superoleophilic PU sponges. For example, Wang et al. [13] prepared a superhydrophobic and flame retardant PU sponge in an aqueous suspension containing alkaline water/ethanol, dopamine, fly ash, and dodecanethiol through the dip-coating method. The modified sponge presented a good oil sorption capacity (34-47 g/g) and separation efficiency for oil-in-water emulsions (93.0%).

To modify PU, graphene-based materials such as reduced graphene oxide (RGO) and graphene oxide (GO) have been utilized, owing to their remarkable physicochemical properties such as low density, large specific surface area, and biocompatibility [14, 15]. Moreover, GO sheets are decorated with hydrophilic functional groups including hydroxyl, epoxy, carbonyl, and carboxylic on their basal planes and edges. The oxygen-containing functional groups of GO can play a main role in modification of PU sponge. Liu et al. [16] synthesized silica NPs on GO surface ( $\text{SiO}_2/\text{GO}$ ). Then, PU sponge was dipped into the ethanol/water solution containing the  $\text{SiO}_2/\text{GO}$  to obtain a hydrophobic  $\text{SiO}_2/\text{GO}/\text{PU}$  (WCA= $145^\circ$ ). The modified PU sponge presented a high sorption capacity of 180 g/g for chloroform and exhibited a high stability for 10 cycles without any change of the intrinsic

structure of the PU sponge. Some of us [15] prepared a flame retardant, 98  
superhydrophobic/superoleophilic graphene-based PU sponge by utilizing a simple *in-situ* 99  
chemical method. The modified sponge displayed a high sorption capacity of 80 g/g for 100  
chloroform and exhibited greater thermal stability with respect to neat PU. 101

By incorporation of magnetic nanoparticles (NPs) into graphene-based PU sponge, both 102  
oil/water separation and fire retardancy properties could be enhanced [17-19]. Moreover, the 103  
whole separation process could be controlled through external stimulation using a magnet. 104  
Yang et al. [20] fabricated graphene aerogels modified by magnetic carbon nanospheres 105  
through one-step hydrothermal *in-situ* electrostatic self-assembling method. The prepared 106  
aerogels showed excellent thermal stability, flame retardancy, and magnetic properties. In 107  
addition, they achieved high sorption capacity (187–537 g/g) for several organic solvents and 108  
oils. Boukherroub et al. [21] prepared a superhydrophobic/superoleophilic PU sponge 109  
through a two-step functionalization process. The modified sponge not only presented high 110  
sorption capacity (40-75 g/g) for several oils, but it was also successfully applied for effective 111  
separation of O/W emulsions. 112

Herein, we prepared  $\text{Mn}_{0.01}\text{Co}_{0.90}\text{Fe}_3\text{O}_4$  nanoparticles (MCFO NPs) *via* a simple 113  
solvothermal method. Then, the prepared NPs were grown on GO to form MCFO/RGO 114  
nanocomposite and incorporated into a PU sponge. Introduction of some magnetic metal ions 115  
such as manganese and cobalt ions can affect the thermal degradation and flame-retardancy 116  
of PU [22]. To the best our knowledge, this is the first demonstration of the preparation of PU 117  
functionalized with MCFO/RGO nanocomposite. The MCFO/RGO/PU sponge displayed 118  
excellent superhydrophobic property with high selectivity and outstanding recyclability for 119  
oil sorption. It could be magnetically directed to adsorb oils and organic solvents up to 131 120  
times of its weight with respect to neat PU sponge. Additionally, one of the great success of 121  
this work is to separate oil droplets from oil-in-water emulsions by the prepared 122

MCFO/RGO/PU sponge. In fact, oil/water emulsions especially in the presence of surfactants	123
are more difficult to separate in real spaces, as compared to immiscible oil/water mixtures.	124
<b>2. Experimental</b>	125
<i>2.1. Preparation of Mn<sub>0.01</sub>Co<sub>0.90</sub>Fe<sub>3</sub>O<sub>4</sub> (MCFO)/RGO nanocomposite</i>	126
<i>In situ</i> growth of MCFO NPs on GO to produce MCFO/RGO nanocomposite was performed	127
by a facile solvothermal method. For this purpose, 80 mg of GO were dispersed in 80 mL of	128
ethylene glycol and 3 mL of ethyl acetate in an ultrasonic bath for 60 min. Afterwards,	129
Mn(CH <sub>3</sub> COO) <sub>2</sub> ·4H <sub>2</sub> O (0.5 mmol), Co(CH <sub>3</sub> COO) <sub>2</sub> ·4H <sub>2</sub> O (0.5 mmol), FeCl <sub>3</sub> ·6H <sub>2</sub> O (2.0 mmol),	130
and CH <sub>3</sub> COONa (1.2 mmol) were added into the aforementioned mixture and sonicated to	131
form a homogeneous slurry. The mixture was sealed in a stainless-steel autoclave and heated	132
at 200 °C for 72 h. The obtained precipitate was collected by centrifugation and washed with	133
Milli-Q water and ethanol, and finally dried in an oven at 50 °C for 24 h to obtain the	134
MCFO/RGO nanocomposite. The similar procedure was used to prepare magnetic	135
MCFO/RGO nanocomposites with different amounts of GO (70, 80, and 90 mg). The	136
resulting nanocomposites are denoted as MCFO/RGO-1 (GO=70 mg), MCFO/RGO-2	137
(GO=80 mg), and MCFO/RGO-3 (GO=90 mg).	138
<i>2.2. Preparation of magnetic MCFO/RGO/PU sponge</i>	139
PU sponge pieces (1×1×1 cm <sup>3</sup> ) were washed first with acetone and then with DI water in an	140
ultrasonic bath for 10 min and dried in an oven for 2 h at 80 °C. The clean sponge was dipped	141
in a solution containing MCFO/RGO-1 (50 mg), ethanol (10 mL), and Milli-Q (10 mL) and	142
maintained under stirring for 6 h at room temperature. After that, the treated sponge was	143
dried by freeze-drying for 5 h. A similar approach was applied for the preparation of	144
MCFO/RGO-2 and 3 samples. Similarly, three sponges denoted as MCFO/RGO/PU-1	145

(GO=70 mg), MCFO/RGO/PU-2 (GO=80 mg) and MCFO/RGO/PU-3 (GO=90 mg) were prepared.

### 2.3. Preparation of RGO/PU sponge

GO (80 mg) was dispersed in water (80 mL) and sealed in a stainless-steel autoclave and heated at 200 °C for 72 h to prepare RGO powder. A clean PU sponge was dipped into the solution containing RGO (50 mg), ethanol (10 mL), and Milli-Q (10 mL) and maintained under stirring for 6 h at room temperature. Finally, the treated sponge was dried by freeze-drying for 5 h.

### 2.4. Oil and solvent sorption

The sorption capacity,  $Q$ , (mass/mass, g/g) of neat PU, RGO/PU, and MCFO/RGO/PU sponges for oils and organic solvents including cooked oil, chloroform, toluene, ethyl acetate, and dichloromethane was measured at room temperature. The sorption capacity was calculated using Eq. 1:

$$Q = \frac{m_1 - m_0}{m_0} \quad (1)$$

where  $m_0$  and  $m_1$  represent respectively the mass of MCFO/RGO/PU sponge before and after sorption.

### 2.5. Emulsion separation using MCFO/RGO/PU sponge

#### 2.5.1. Surfactant-free water-in-oil emulsion (W/O)

Water (1 mL) was added to toluene (9 mL) and sonicated for 1 h to produce a milky white solution. The MCFO/RGO/PU sponge ( $1 \times 1 \times 1 \text{ cm}^3$ ) was immersed in 4 mL of the emulsion for 30 min till the emulsion became colorless. The content of toluene in water was determined using a calibration curve generated from the UV-vis absorbance spectra at 260



nm of different concentrations of toluene dispersed in methanol/water (3/1, v/v) solution (see SI). 168

2.5.2. Surfactant-stabilized oil-in-water emulsion (O/W) 169

Pluronic F-127 (0.1 mg/mL in water) was added to toluene/water (1:9, v/v) and sonicated for 171

1 h to prepare a milky white solution. The MCFO/RGO/PU sponge (1×1×1 cm<sup>3</sup>) was 172

immersed in 4 mL of the emulsion for 30 min till the emulsion became colorless. The content 173

of toluene in water was determined as above. 174

3. Results and discussion 175

3.1. Preparation of magnetic superhydrophobic/superoleophilic (MCFO/RGO/PU) sponge 176

MCFO/RGO/PU sponge was fabricated through a two-step process including a 178

solvothermal method to coat MCFO NPs on RGO nanosheets, followed by dip-coating 179

method to load MCFO/RGO nanocomposite onto the PU surface. The preparation process of 180

MCFO/RGO nanocomposite is illustrated in **Figure S1**. The well-dispersed GO nanosheets in 181

ethylene glycol and ethyl acetate solution served as a substrate for the nucleation and *in situ* 182

growth of MCFO NPs [23]. Co<sup>2+</sup>, Mn<sup>2+</sup>, and Fe<sup>3+</sup> ions can be anchored on the GO surface 183

during the nucleation process through electrostatic interactions between the positively 184

charged metal ions and the negatively charged GO sheets [24, 25]. By elapsing time, GO 185

was partially reduced to RGO and NPs were formed around the heterogeneous nucleation 186

sites on the RGO sheets [26]. The solvents may act as structure-directing agents and prevent 187

agglomeration of the NPs during the nucleation step. Finally, MCFO/RGO/PU 188

nanocomposite was prepared by immersion of a PU sponge into a MCFO/RGO colloidal 189

solution. ~~The sponge and nanocomposite were connected through hydrogen bonding between~~ 190

~~HNCO groups of PU and the deprotonated oxygen functional groups of MCFO/RGO. In~~ 191

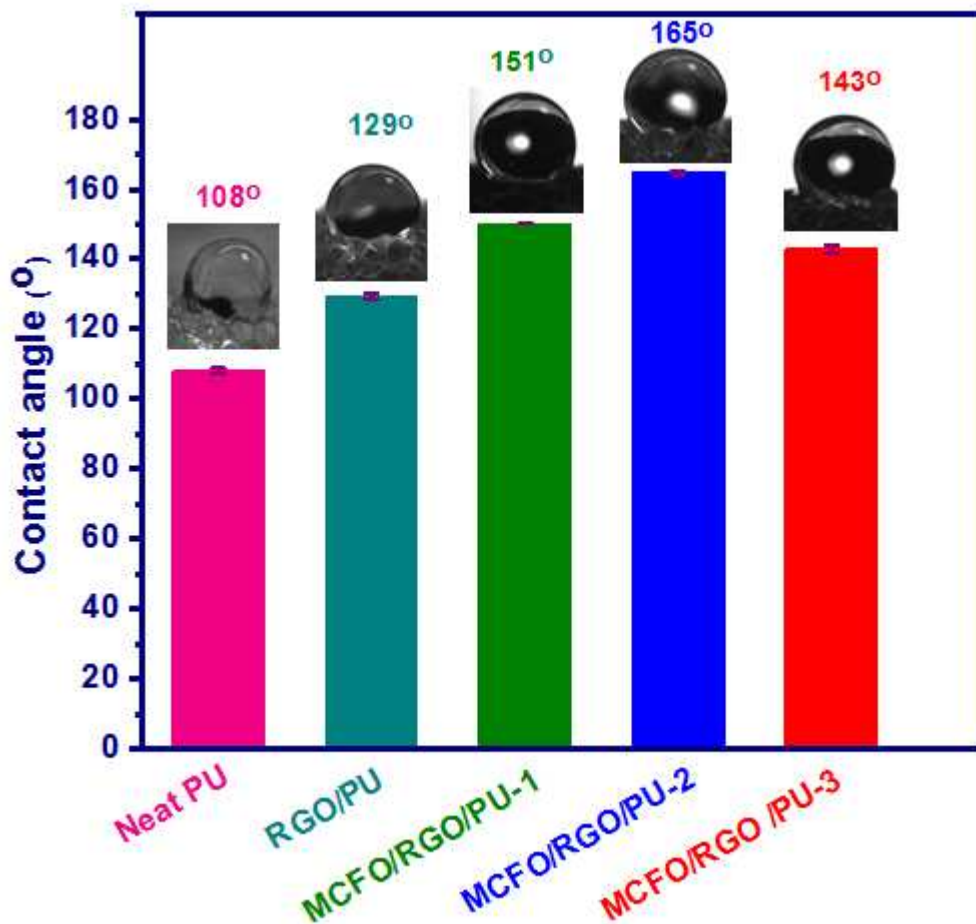
~~addition,  $\pi - \pi$  stacking interactions between RGO sheets and aromatic groups of PU sponge~~ 192

could strengthen the bonding of MCFO/RGO nanocomposite on the PU surface. The MCFO 193  
nanoparticles are strongly immobilized on the RGO surface owing to the presence of plenty 194  
of oxygen-containing groups on the GO precursor. The MCFO/RGO is in turn adsorbed on 195  
PU surface through  $\pi$ - $\pi$  stacking interactions between RGO sheets and aromatic groups of 196  
the PU sponge. 197

### 3.2. *Characterization* 198

#### 3.2.1. *WCA measurements* 199

To investigate the wetting properties of neat PU and prepared sponges including 200  
RGO/PU, MCFO/RGO/PU-1, MCFO/RGO/PU-2, and MCFO/RGO/PU-3, the WCA was 201  
measured on three different locations on the surface of the sponges using 4  $\mu$ L of DI water 202  
droplets at room temperature. As shown in **Figure 1**, neat PU sponge displays a WCA of 108 203  
 $\pm 1^\circ$ , which increased to  $129 \pm 1^\circ$  after RGO coating. The hydrophobicity of the PU sponge 204  
was enhanced upon loading MCFO/RGO on its surface, as witnessed by the apparent WCA 205  
values of MCFO/RGO/PU-1 ( $151 \pm 1^\circ$ ), MCFO/RGO/PU-2 ( $165 \pm 1^\circ$ ) and MCFO/RGO/PU- 206  
3 ( $143 \pm 1^\circ$ ). It is obvious that the WCA of MCFO/RGO/PU-2 herein after named as 207  
MCFO/RGO/PU is greater than those of other modified sponges. By increasing the amount of 208  
GO above 80 mg, the WCA decreased. It is believed that higher amounts of GO can fill the 209  
pores of the PU sponge and decrease its surface roughness and wettability. Also, it is likely 210  
that RGO nanosheets are aggregated with increasing the amount of GO precursor. 211



**Figure 1.** WCA values of neat PU and different MCFO/RGO/PU sponges.

To investigate the oleophilicity of MCFO/RGO/PU sponge, cooked oil was used. When cooked oil was dropped on the surface of the sponge, the oil droplet immediately spread out and meanwhile permeated into the porous structure of the sponge. Therefore, no contact angle could be measured, demonstrating the superoleophilicity of the MCFO/RGO/PU sponge (**Movie S1**). Owing to the superhydrophobicity and superoleophilicity of MCFO/RGO/PU sponge, it has a great potential for practical application in oil/water separation processes.

### 3.2.2. XRD, FTIR, Raman, XPS, and TEM analyses

The XRD patterns of GO and MCFO/RGO are depicted in **Figure S2**. The sharp peak at  $2\theta=10.85^\circ$  in the XRD pattern of GO corresponds to the (002) crystalline plane with a large

interlayer distance of 8.2 Å due to the presence of oxygen containing functional groups. This peak disappears for MCFO/RGO nanocomposite, confirming the partial recovery of graphitic crystal structure. Indeed, two new peaks at 24.36° and 42.98° assigned respectively to the (002) and (102) crystalline planes of RGO appeared. The XRD pattern of MCFO/RGO comprises six characteristic diffraction peaks at  $2\theta=18.48^\circ$ ,  $30.06^\circ$ ,  $35.55^\circ$ ,  $53.86^\circ$ ,  $56.95^\circ$ , and  $62.67^\circ$  attributed respectively to the (111), (220), (311), (422), (511), and (440) lattice planes. These peaks indicate magnetite  $\text{Fe}_3\text{O}_4$  phase formation (JCPDS-96-900-9770) [27]. The two peaks at  $2\theta=44.50^\circ$  and  $47.65^\circ$  correspond respectively to the (002) and (101) crystalline planes of hcp-Co (JCPDS-96-901-0968) [28]. In addition, the peak at about  $2\theta=42.98^\circ$  (400) (JCPDS-96-153-9040) is ascribed to cubic-Mn [29]. The atomic ratio of Mn and Co was calculated by ICP analysis. The presence of magnetite  $\text{Fe}_3\text{O}_4$  phase and the atomic ratio of Co and Mn elements (0.90:0.01), determined from ICP analysis, suggest the formation of  $\text{Mn}_{0.01}\text{Co}_{0.90}\text{Fe}_3\text{O}_4$ .

The FTIR spectra of GO, RGO, and magnetic MCFO/RGO are displayed in **Figure S3**. The FTIR spectrum of GO shows bands at 3381, 1726, 1620, 1383, 1228, and  $1054\text{ cm}^{-1}$  ascribed respectively to O-H, C=O carboxyl, C=C in the aromatic ring, C-OH, epoxy C-O, and alkoxy C-O vibrations. The FTIR spectrum of RGO exhibits four weak bands at 3381, 1726, 1582, and  $1052\text{ cm}^{-1}$  assigned respectively to O-H, C=O carboxyl, C=C, and alkoxy C-O vibration modes [15]. The C=C stretching mode of RGO was blue-shifted from 1620 to  $1582\text{ cm}^{-1}$ , suggesting an increased structural ordering and restoration of  $\text{sp}^2$  domains due to hydrothermal reduction of GO [30]. The FTIR spectrum of MCFO/RGO is composed of bands at 3381, 1620, 1390, and  $680\text{ cm}^{-1}$  assigned to vibrations of O-H, C=C in the aromatic ring, Fe-O stretching mode of  $\text{Fe}_3\text{O}_4$  [31], and Co(III)-O, respectively [32]. The bands at 1040, 1172, and  $1230\text{ cm}^{-1}$  are related to the interactions among Mn, OH, and O groups atoms/groups [33, 34].

**Figure S4** displays the Raman spectrum of MCFO/RGO with four obvious peaks of RGO, 250  
marked as D ( $1354\text{ cm}^{-1}$ ), G ( $1586\text{ cm}^{-1}$ ), 2D ( $2700\text{ cm}^{-1}$ ), and D+G ( $2950\text{ cm}^{-1}$ ) bands [35]. 251  
The G band is related to  $E_{2g}$  vibration mode of  $sp^2$  carbon domains, whereas the D band 252  
arises from structural defects [36]. The 2D and D+G bands are due to the overtone of the D 253  
band and the combination mode induced by disorder, respectively [37]. Also, the peaks at 254  
around  $663$  and  $672\text{ cm}^{-1}$  are characteristic of manganese oxide [38] and  $\text{Fe}_3\text{O}_4$  NPs [39], 255  
respectively. It is likely that both peaks are overlapped. The peak at  $811\text{ cm}^{-1}$  is attributed to 256  
cobalt oxide [40]. There is an extra peak at  $960\text{ cm}^{-1}$  due to silicon [41]. 257

To assess the chemical composition, XPS analysis of MCFO/RGO nanocomposite was 258  
performed. **Figure S5a** depicts the XPS survey spectrum of MCFO/RGO, comprising peaks 259  
due to C 1s ( $285\text{ eV}$ ), O 1s ( $532\text{ eV}$ ), Mn 2p ( $642\text{ eV}$ ), Fe 2p ( $710\text{ eV}$ ), and Co 2p ( $781\text{ eV}$ ), 260  
which is in a good agreement with the composition of the material. The atomic percentage of 261  
C 1s, O 1s, Co 2p, Fe 2p, and Mn 2p elements are 52.52, 32.75, 10.47, 3.77, and 0.49 at.%, 262  
respectively. The C 1s core level spectrum can be deconvoluted into four peaks at binding 263  
energies of 284.2, 285.1, 286.2, and 288.4 eV due to C=C,  $sp^3$  C-C, epoxy and/or hydroxyl 264  
(C-O/C-OH), and carboxyl (O-C=O) functional groups, respectively (**Figure S5b**). The O 1s 265  
spectrum was curve-fitted with two peaks at binding energies of 531.5 and 532.8 eV assigned 266  
to epoxy (C-O) and hydroxyl (C-OH) groups, respectively (**Figure S5c**). As shown in **Figure** 267  
**S5d**, the high-resolution spectrum of the Mn 2p contains a pair of peaks located at about 268  
 $642.2$  and  $653.6\text{ eV}$  attributed respectively to Mn  $2p_{3/2}$  and Mn  $2p_{1/2}$ . The Mn  $2p_{3/2}$  peak is 269  
deconvoluted into three peaks at  $640.4$ ,  $642.1$ , and  $643.5\text{ eV}$  related to  $\text{Mn}^{2+}$ ,  $\text{Mn}^{3+}$ , and  $\text{Mn}^{4+}$  270  
valence states, respectively [42]. The Mn  $2p_{1/2}$  peak can be fitted with three peaks at binding 271  
energies of  $651.1$ ,  $652.7$ , and  $654.1\text{ eV}$  due to  $\text{Mn}^{2+}$ ,  $\text{Mn}^{3+}$  and  $\text{Mn}^{4+}$  valence states, 272  
respectively [42-44]. Also, two satellites are observed at  $646.1$  and  $659.3\text{ eV}$  [42, 45]. The 273  
Mn peak in the survey spectrum is weak, indicating that the content of Mn ions in 274

MCFO/RGO is low and Mn ions are doped into the interior of the nanocomposite. **Figure S5e** illustrates the Fe 2p XPS core level spectrum, which is deconvoluted in two main peaks including Fe 2p<sub>3/2</sub> and Fe 2p<sub>1/2</sub>. Fe 2p<sub>3/2</sub> can be fitted with three peaks at binding energies of 705.2, 709.9, and 712.1 eV which are the characteristic peaks of Fe<sup>0</sup>, Fe<sup>2+</sup>, and Fe<sup>3+</sup>, respectively [46]. Fe 2p<sub>1/2</sub> peak can be fitted with two peaks at 724.0 and 726.2 eV ascribed to Fe<sup>2+</sup> and Fe<sup>3+</sup>, respectively [47]. Moreover, two satellite peaks are observed at 719.1 and 733.1 eV which are due to Fe<sup>3+</sup> and Fe<sup>2+</sup> chemical states present in Fe<sub>3</sub>O<sub>4</sub> phase [48]. The high-resolution spectrum of the Co 2p in **Figure S5f** shows characteristic peaks of Co at about 781.3 eV (Co 2p<sub>3/2</sub>) and 796.1 eV (Co 2p<sub>1/2</sub>). The Co 2p<sub>3/2</sub> spectrum can be deconvoluted into two peaks at 781.0 and 783.6 eV that are the characteristic peaks for Co<sup>3+</sup> 2p<sub>3/2</sub> and Co<sup>2+</sup> 2p<sub>3/2</sub>, respectively. In addition, the deconvolution of Co 2p<sub>1/2</sub> shows two peaks at 796.2 and 798.4 eV which correspond to Co<sup>3+</sup> 2p<sub>1/2</sub> and Co<sup>3+</sup> 2p<sub>1/2</sub>, respectively. Also, two satellites are observed at 787.0 and 803.2 eV [49].

The TEM images of GO and MCFO/RGO nanocomposite are exhibited in **Figure S6**. The TEM of GO displays a characteristic crumpled surface. The TEM of MCFO/RGO shows the presence of MCFO nanospheres with a size of 3-6 nm attached onto RGO sheets.

### 3.2.3. XRD and FTIR of MCFO/RGO/PU sponge

The XRD patterns of neat PU and MCFO/RGO/PU sponges are depicted in **Figure S7**. The peak at  $2\theta=20.7^\circ$  in the XRD pattern of neat PU corresponds to the (110) crystalline plane [50]. In addition, the characteristic peaks of MCFO/RGO nanocomposite and PU sponges are clearly observed in MCFO/RGO/PU sponge XRD pattern. A very strong broad amorphous peak at about  $2\theta=20.7^\circ$  is ascribed to the (110) plane and related to the disordered structure of the PU matrix amorphous phase [51]. Besides, the observed diffraction peaks at  $2\theta$  values of  $18.48^\circ$ ,  $24.36^\circ$ ,  $30.06^\circ$ ,  $35.55^\circ$ ,  $42.98^\circ$ ,  $44.50^\circ$ ,  $47.65^\circ$ ,  $53.6^\circ$ ,  $56.95^\circ$ , and  $62.67^\circ$  corresponding respectively to the lattice planes of (111), (002), (220), (311), (400), (002),

(101), (422), (511), and (440) indicate that the MCFO/RGO nanocomposite was successfully loaded onto the PU sponge surface.

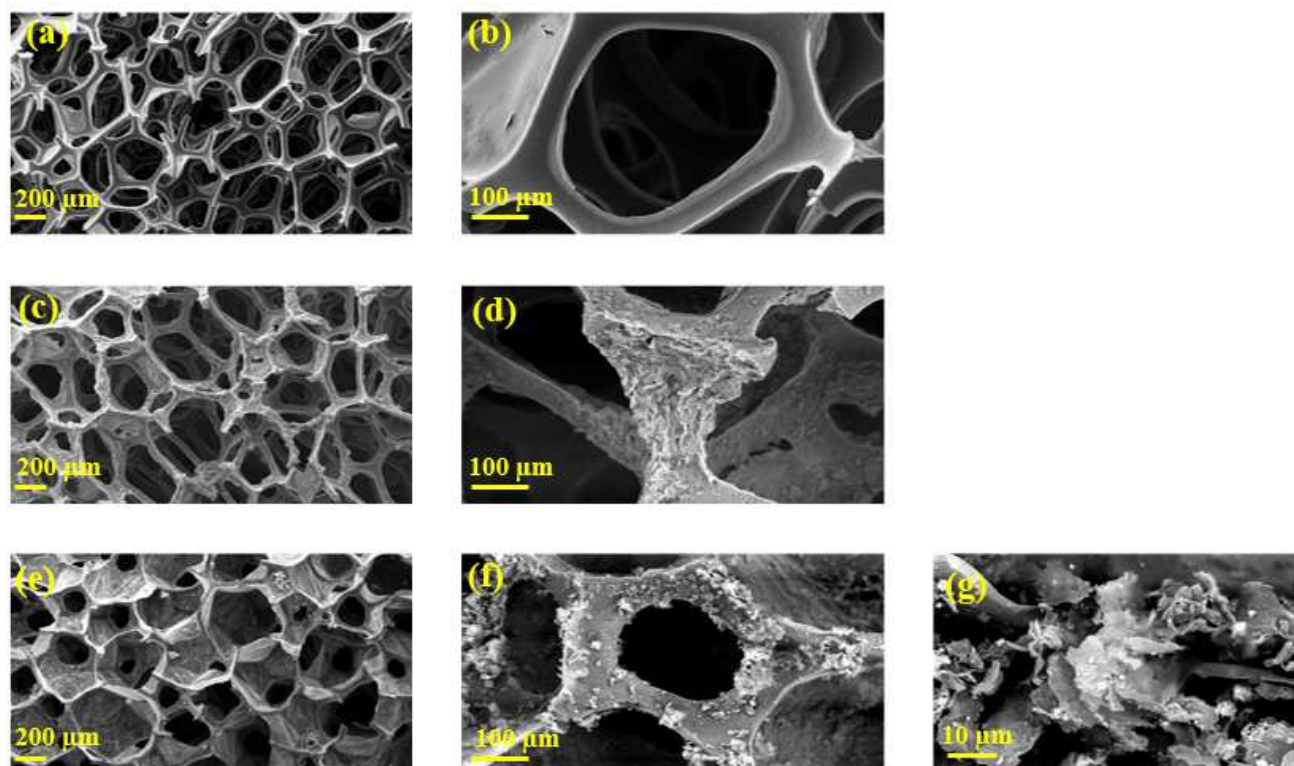
The FTIR spectra of neat PU and MCFO/RGO/PU sponges are displayed in **Figure S8**. The neat PU sponge shows some distinguished bands at  $3270\text{ cm}^{-1}$  ( $\nu_{\text{N-H}}$ ,  $\nu_{\text{O-H}}$ ),  $2850\text{--}2970$  ( $\nu_{\text{C-H}}$ ,  $-\text{CH}_3$  and  $-\text{CH}_2$ ),  $2271\text{ cm}^{-1}$  ( $\nu_{\text{as-NCO}}$ ),  $1722\text{ cm}^{-1}$  ( $\nu_{\text{C=O}}$  in the amide),  $1637\text{ cm}^{-1}$  ( $\nu_{\text{C=O}}$  in the urea),  $1529\text{ cm}^{-1}$  ( $\nu_{\text{C=C}}$ ),  $1445\text{ cm}^{-1}$  (C–H deformation),  $1072\text{ cm}^{-1}$  ( $\nu_{\text{C-O-C}}$ ), and  $671\text{ cm}^{-1}$  (bending vibration of C–H) [52]. The FTIR spectrum of MCFO/RGO/PU sponge displays the same characteristic bands and positions like neat PU except for -NCO group. PU could form chemical bonds with materials through the NCO group of isocyanates. The absence of -NCO group in the FTIR spectrum of the modified PU sponge obviously demonstrates that the MCFO/RGO nanocomposite could be anchored on the PU surface effectively.

#### 3.2.4. SEM images of neat PU, RGO/PU, MCFO/RGO/PU sponges

The SEM images of neat PU, RGO/PU, and magnetic MCFO/RGO/PU sponges are displayed in **Figure 2**. As shown in **Figure 2a**, the interconnected 3D porous structure of neat PU with a pore size ranging  $100\text{--}500\text{ }\mu\text{m}$  makes it appropriate for sorption and storage of oil pollutants. Also, the smooth (absence of wrinkles) surface of the neat PU is obvious in its high magnification image (**Figure 2b**). As displayed in **Figure 2c** and **2d**, the surface of RGO/PU sponge exhibits higher roughness with respect to that of neat PU. From the SEM images in **Figure 2e** and **2f**, it is obvious that the MCFO/RGO/PU sponge maintains its porous structure with numerous tiny pores with respect to neat PU. It makes it appropriate for emulsion breakage. From the enlarged image of modified PU (**Figure 2g**), numerous irregular particles can be observed on the surface compared with the neat PU. This results in an enhancement of surface roughness and increase of superhydrophobic/superoleophilic property. In addition, compared with neat PU and RGO/PU sponges, such hierarchical

roughness could effectively reduce the diameter of sponge pores to the diameter of 325  
toluene/water emulsion droplets' size. 326

The EDS spectrum of MCFO/RGO/PU sponge in **Figure S9** demonstrates the presence of 327  
Co, Fe, Mn, O, and C elements in the modified sponge in good agreement with the XPS 328  
analysis. 329



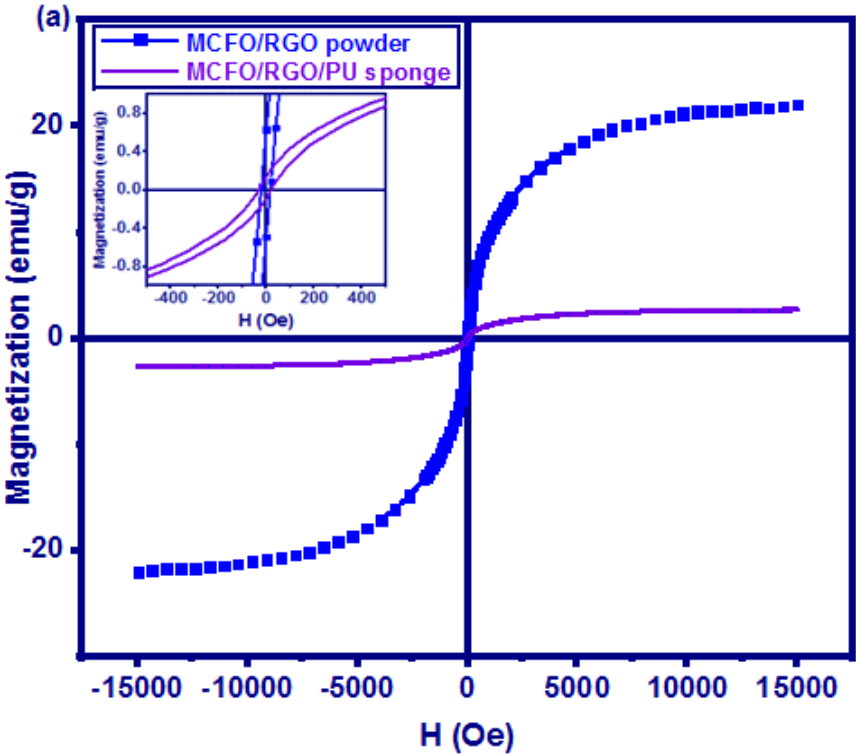
**Figure 2.** SEM images of (a-b) neat PU, (c-d) RGO/PU, and (e-g) MCFO/RGO/PU 331  
sponges. 332

### *Magnetic properties* 333

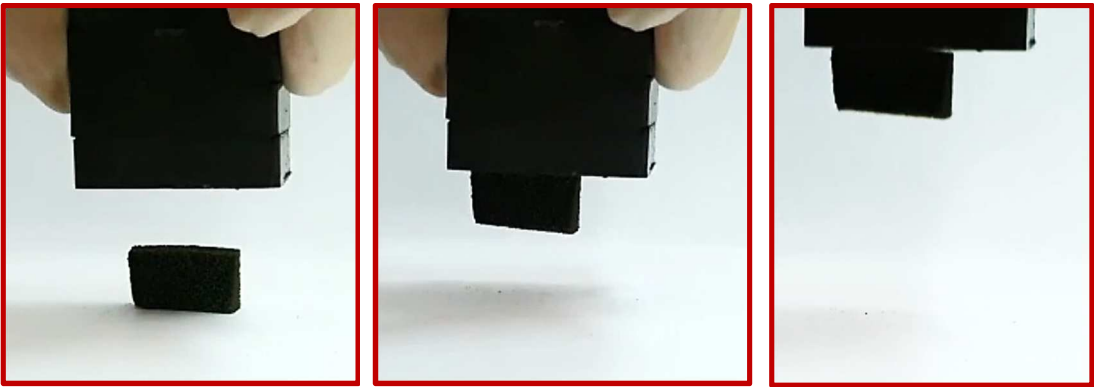
The magnetic curves of the MCFO/RGO nanocomposite and MCFO/RGO/PU sponge are 334  
depicted in **Figure 3a**. The MCFO/RGO nanocomposite is ferromagnetic with remnant 335  
magnetization,  $M_r$ , coercivity,  $H_c$ , and specific saturation magnetization,  $M_s$ , of  $0.63 \text{ emu g}^{-1}$ , 336  
 $20.0 \text{ Oe}$ , and  $22.05 \text{ emu g}^{-1}$ , respectively. Also, the MCFO/RGO/PU sponge exhibited 337  
ferromagnetic property with  $M_r$ ,  $H_c$ , and  $M_s$  of  $0.14 \text{ emu g}^{-1}$ ,  $30.0 \text{ Oe}$ , and  $2.72 \text{ emu g}^{-1}$ , 338  
respectively. It is obvious that the magnetic property of MCFO/RGO/PU sponge is decreased 339



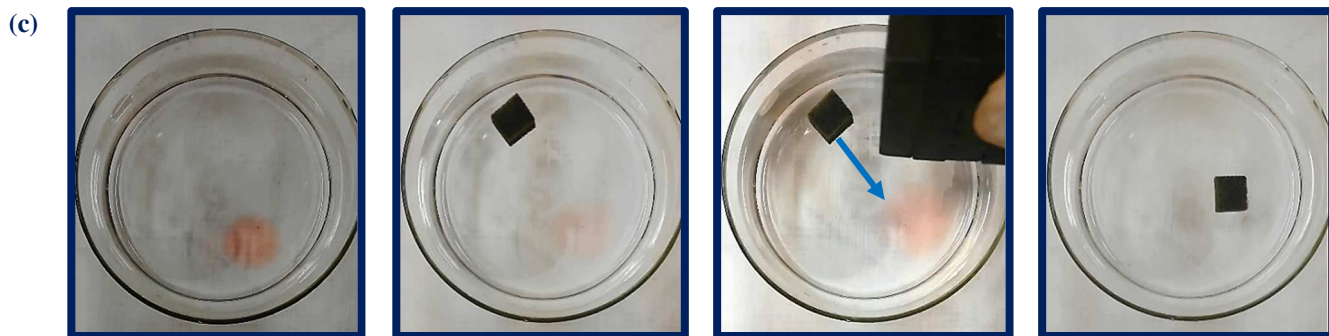
compared to that of MCFO/RGO nanocomposite due to the presence of relatively lower 340  
magnetic content. Nevertheless, the magnetic properties of MCFO/RGO/PU sponge are 341  
sufficient to manipulate the sponge using a magnet, as illustrated in **Figure 3b** and **movie S2**. 342  
**Figure 3c** and **movie S3** show the magnetic response of MCFO/RGO/PU sponge. The 343  
modified sponge can be easily driven to adsorb droplet of contaminant by using a magnet. 344



(b)



345



**Figure 3.** (a) Magnetization curves of MCFO/RGO nanocomposite and MCFO/RGO/PU 346  
sponge, (b) lifting the modified sponge *via* a magnet bar, and (c) the removal process of 347  
hydraulic oil by a piece of prepared sponge under a magnetic field. 348

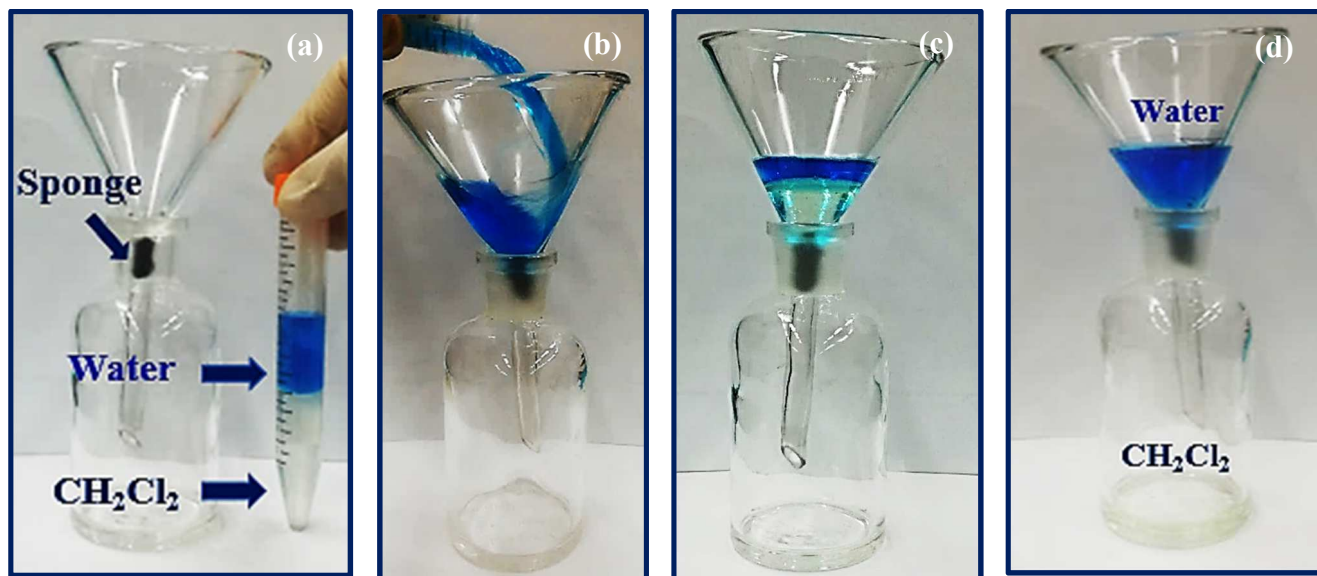
### 3.4. MCFO/RGO/PU sponge stability in harsh conditions 349 350

Sorbent stability is an important factor in practical applications. **Figure S10a** confirms the 351  
excellent stability of the magnetic MCFO/RGO/PU sponge in harsh conditions including 352  
acidic, basic, and saline solutions upon immersion for 6 h at room temperature. In addition, 353  
the excellent stability property of the modified sponge was examined in a wide range of 354  
temperature from -4 to +140 °C. As shown in **Figure S10b**, the WCA is not significantly 355  
impacted by increasing the temperature except at 140 °C where a slight decrease in the WCA 356  
was observed (WCA=159±1°). Therefore, the results prove the chemical and thermal 357  
stabilities of the modified MCFO/RGO/PU sponge, which guarantee its great potential for 358  
practical applications. 359

### 3.5. Separation of oil/water mixtures using the MCFO/RGO/PU sponge 360

In order to demonstrate the oil/water separation performance of the magnetic 361  
MCFO/RGO/PU sponge, a simple separation experiment was designed. The modified sponge 362  
was embedded in a funnel and a mixture of methylene blue-dyed water and dichloromethane 363  
was prepared and poured into the funnel (**Figure 4**). Although CH<sub>2</sub>Cl<sub>2</sub> was successfully 364

separated from water by the sponge, organic solvents with lower density than water could not be easily separated. 365  
366  
367

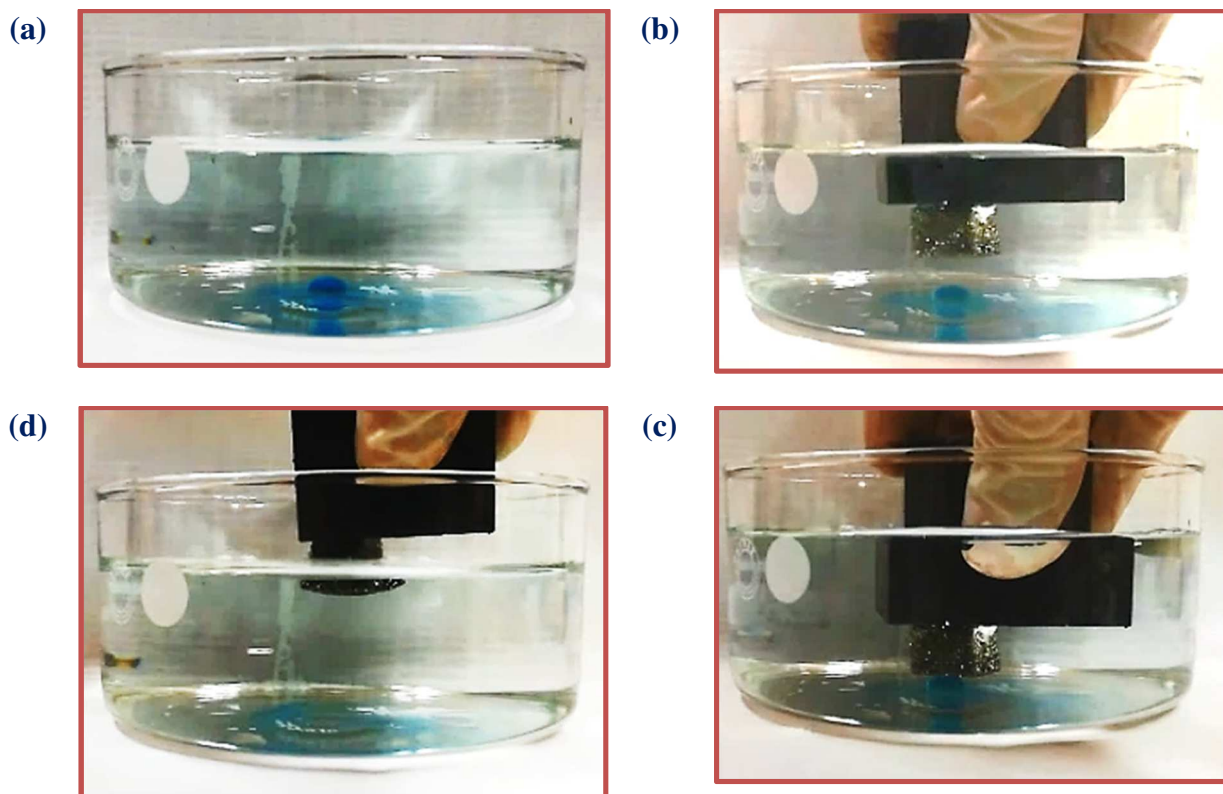


**Figure 4.** Snapshots of dichloromethane and colored water with methylene blue separation by MCFO/RGO/PU sponge. 368  
369

The **movie S4** shows that the dichloromethane permeates through the MCFO/RGO/PU sponge rapidly, whereas the dyed water is maintained on top of the sponge, suggesting that the modified sponge has an excellent oil/water separation performance. 370  
371  
372

### 3.6. *Magnetic property maintenance of MCFO/RGO/PU sponge underwater* 373

To reveal the capability of magnetic response and oil/water separation of the MCFO/RGO/PU sponge, a methylene blue-dyed dichloromethane droplet is used as an oil spill underwater. As illustrated in **Figure 5**, the dichloromethane spill underwater could be rapidly adsorbed through an efficient magnetic drive by MCFO/RGO/PU sponge. Besides, the **movie S5** highlights that the functionalized sponge could effectively separate the oil droplet underwater and be handled by the magnet after oil sorption. 374  
375  
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**Figure 5.** Snapshots of removal of methylene blue dyed underwater dichloromethane droplet 380  
by the MCFO/RGO/PU sponge using a magnet. 381

382

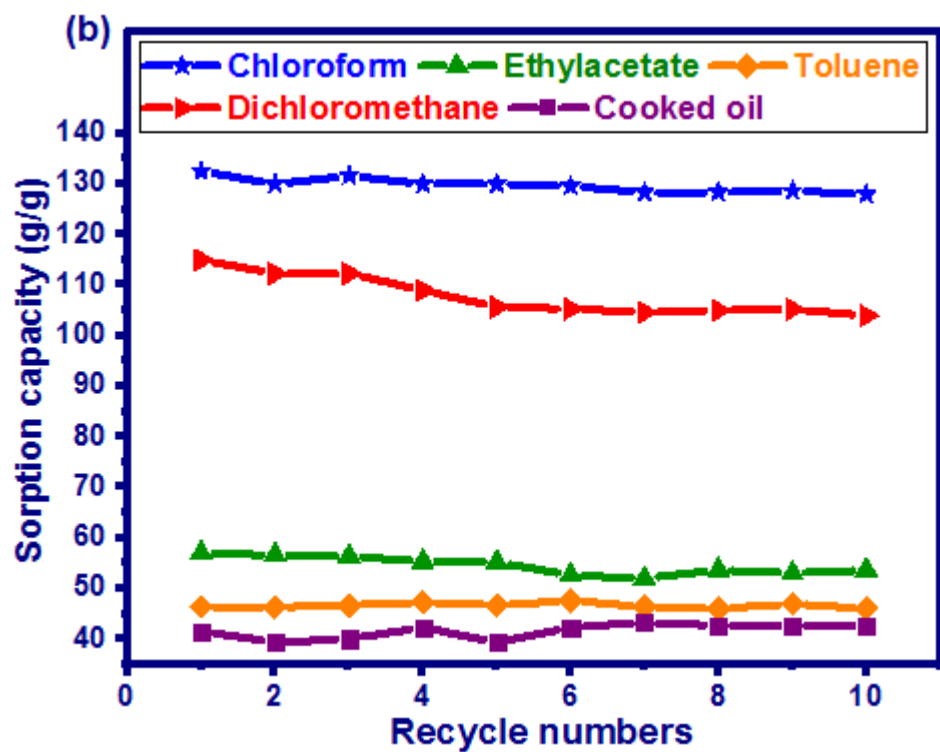
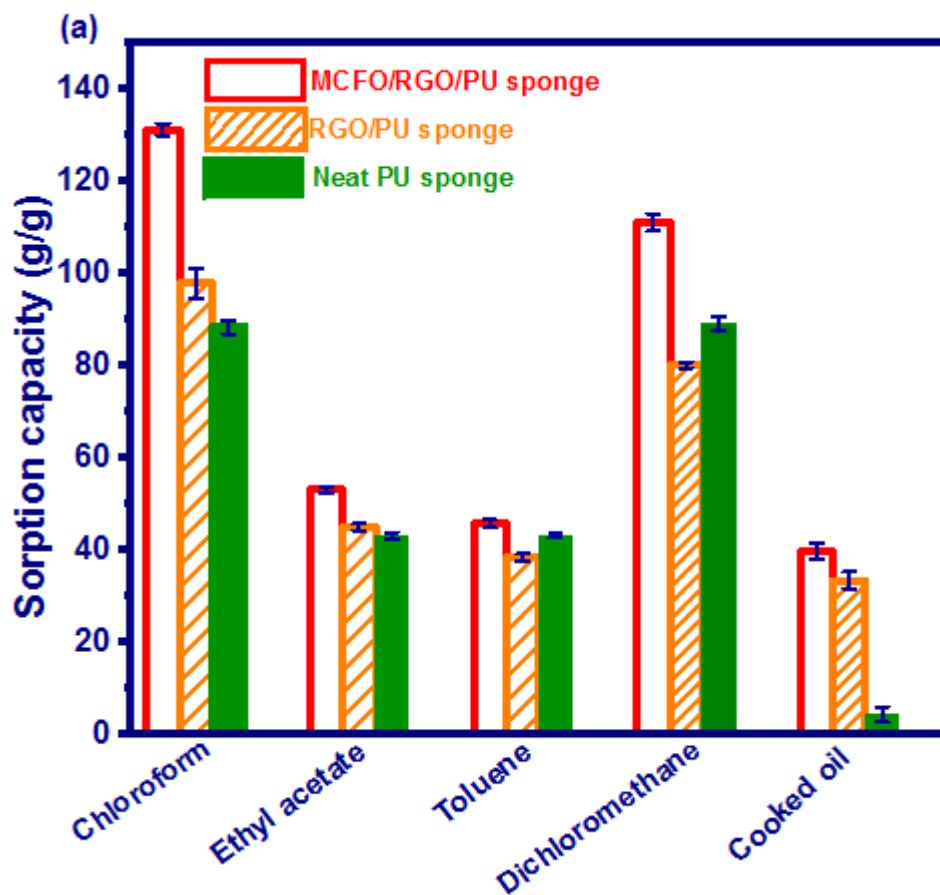
### 3.7. Sorption capacity and recyclability 383

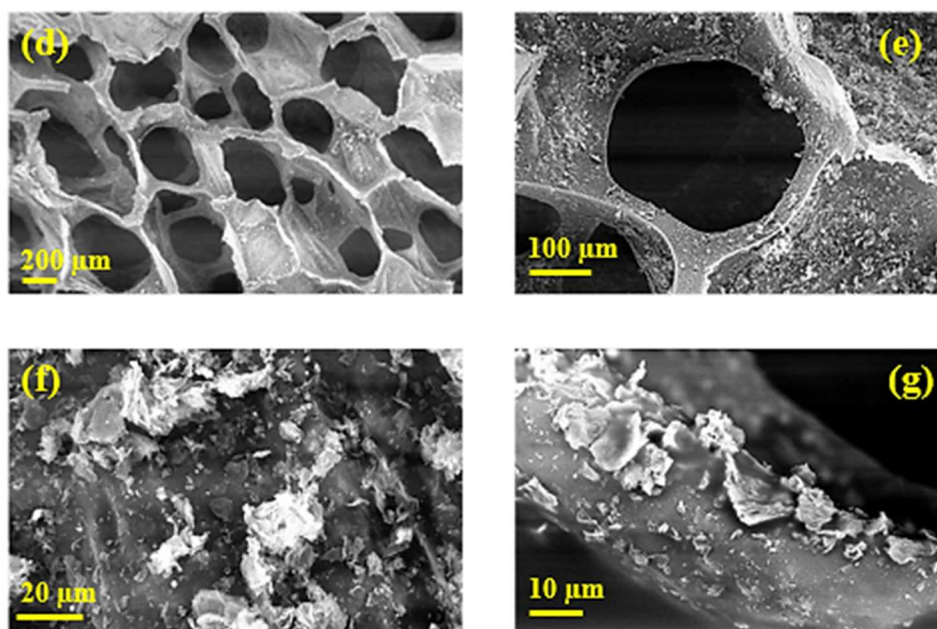
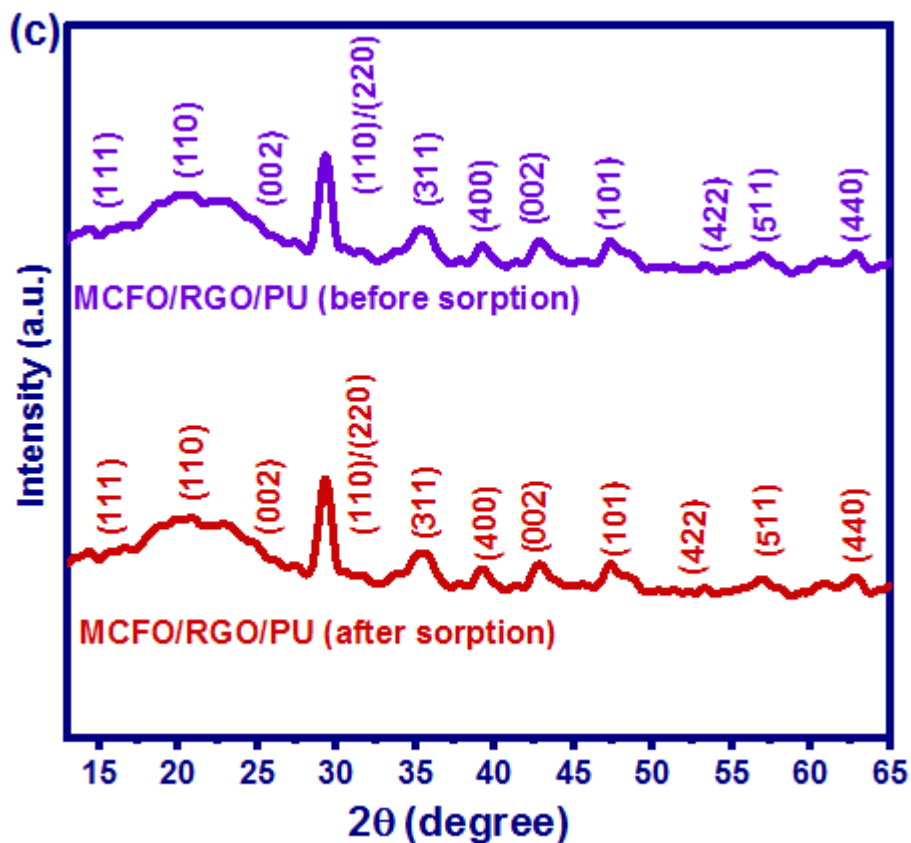
The sorption capacity of neat PU, RGO/PU, and magnetic MCFO/RGO/PU sponges for 384  
oil and organic solvents including cooked oil, chloroform, ethyl acetate, toluene, and 385  
dichloromethane was determined to be in the range of 4.4-88.5, 33.5-98.0, and 39.8-131.4 386  
g/g, respectively (**Figure 6a**). Hence, MCFO/RGO/PU sponge exhibits the best performance 387  
for sorption of oil and organic solvents due to the presence of several active sites on its 388  
surface. The variation of the sorption capacity of the MCFO/RGO/PU sponge for organic 389  
solvents and oils depends on their viscosity and density. The sorption capacity of RGO/PU 390  
for oils and organic solvents is greater than that of neat PU, whereas it is lower than that of 391  
MCFO/RGO/PU. Although RGO sheets enhance the hydrophobicity and oleophilicity of neat 392  
PU, at the same time, they fill some pores of PU. The performance of MCFO/RGO/PU 393  
sponge in comparison with other magnetic PU sponges is provided in **Table S1**. The 394

MCFO/RGO/PU prepared in this work has the greatest sorption capacity, which makes it as a promising sorbent for practical oil separation.

The recyclability of MCFO/RGO/PU sponge for five oil/organic solvents is investigated in **Figure 6b**. A simple mechanical squeezing method was used to desorb oils/organic solvents. The results showed that the MCFO/RGO/PU sponge could be recovered effectively, owing to its excellent mechanical property. The sorption capacity of MCFO/RGO/PU sponge did not change noticeably even after 10 sorption/desorption cycles (**Figure 6b**), which indicates its good structural stability.

XRD and SEM analysis of the MCFO/RGO/PU sponge after ten cycles of toluene sorption were performed. The XRD patterns of MCFO/RGO/PU sponge before and after ten sorption cycles of toluene are depicted in **Figure 6c**. The characteristic peaks of MCFO/RGO/PU sponge are clearly observed after ten sorption cycles, confirming the high stability of the sorbent. Besides, From the SEM images in **Figure 6d**, it is obvious that the MCFO/RGO/PU sponge maintained its porous structure after ten sorption cycles of toluene. In addition, from the enlarged images of the modified PU (**Figure 6e-g**), numerous irregular particles can be observed on the surface, indicating that the MCFO/RGO coating was not released from the PU surface after ten sorption cycles.



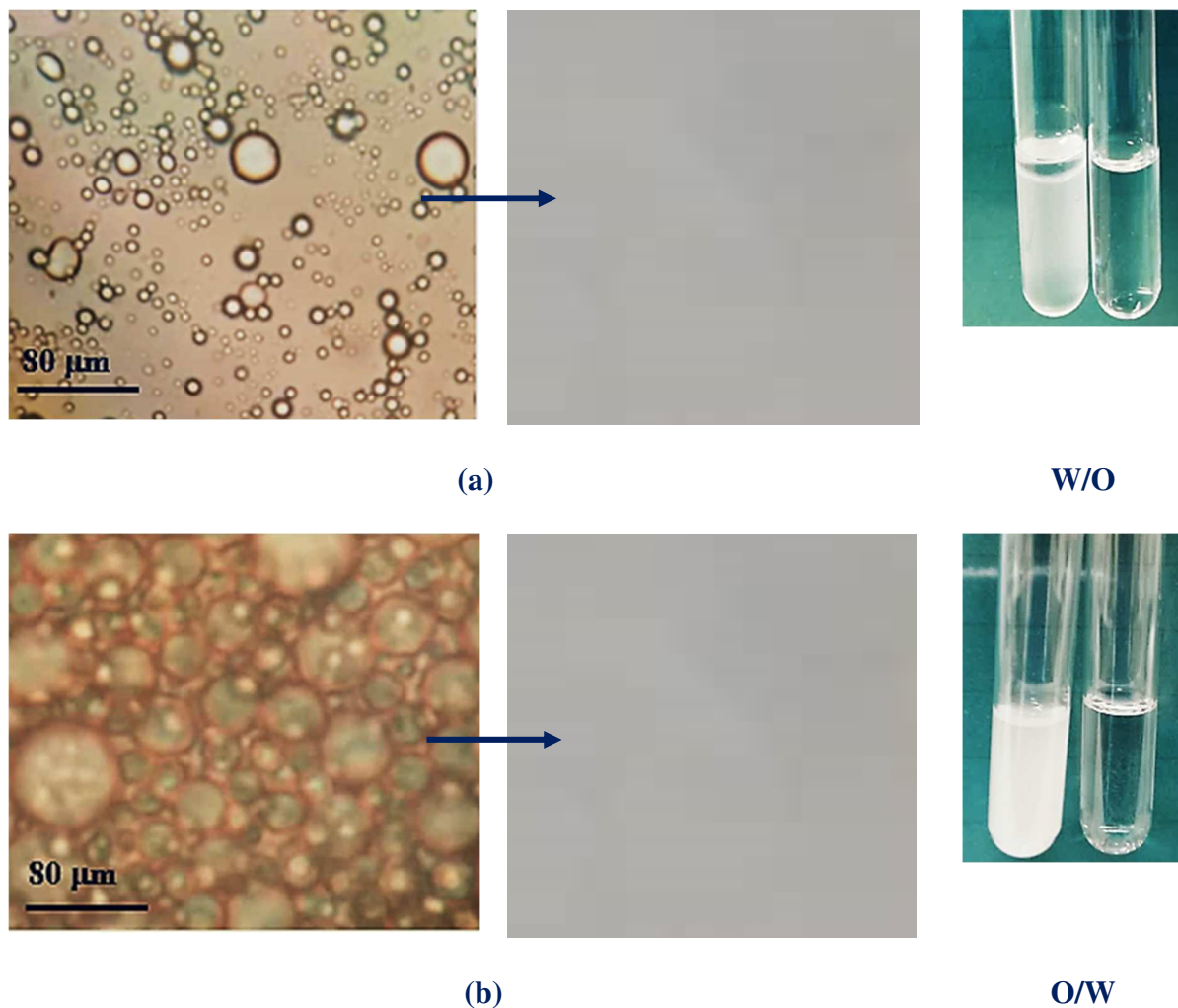


**Figure 6.** (a) The adsorption capacity of PU, RGO/PU, and MCFO/RGO/PU sponges for organic solvents and oils, (b) Recyclability of MCFO/RGO/PU for 10 adsorption cycles, (c) XRD patterns of MCFO/RGO/PU sponge before and after ten sorption cycles of toluene, and (d-g) SEM images of MCFO/RGO/PU sponge after ten sorption cycles of toluene.

### 3.8. Demulsification capacity of MCFO/RGO/PU sponge

The oil sewage, according to their diameter,  $d$ , can be classified into immiscible mixtures, unstable dispersions, and stable emulsions with  $d < 20 \mu\text{m}$ . The size mismatch between porous sponge (100-500  $\mu\text{m}$ ) and emulsified droplets limits the separation efficiency of sponges [1]. It is important to address this challenge while maintaining the high porosity of the sponges [19]. Herein, our strategy is to construct MCFO/RGO decorated PU surface, while preserving its intrinsic porous structure. As shown in SEM images (**Figure 2**), the MCFO/RGO nanocomposite could effectively reduce and match the diameter of sponge pores with that of toluene droplets in toluene/water emulsion. The MCFO/RGO/PU sponge could fully contact with emulsified water droplets, adsorb the toluene, and speed up the water coalescence. To investigate the demulsification ability of the MCFO/RGO/PU sponge, a W/O mixture was prepared by mixing 9 mL of toluene and 1 mL of water and sonicated for 1 h to produce a milky white colloid. In addition, O/W emulsion was prepared by addition of Pluronic F-127 (0.1 mg/mL in water) to toluene/water (1:9, v/v) and subsequently the mixture was emulsified under an ultrasonic bath for 1 h to prepare a milky white colloid. The MCFO/RGO/PU sponge was immersed in 4 mL emulsion until the emulsion became colorless. Before separation, the feeds of W/O and O/W mixtures were like a milky white liquid, as shown in **Figure 7a, b** left. After separation, the collected solutions were transparent compared to the feeds. The toluene content in the water phase was determined by UV-vis spectrophotometry (calibration curve in **Figure S11**). By immersion of the MCFO/RGO/PU sponge, a reduction in the toluene content (W/O) from 90% to 0.1% as well as (O/W) from 10% to 0.01% corresponding to toluene removal of 99.8% and 99.9%, respectively was achieved. It is worth to state that MCFO/RGO/PU sponge not only can separate the immiscible oil/water mixtures effectively, but it is also useful for emulsified oil/water mixtures especially in surfactant-stabilized emulsions.





447 **Figure 7.** Optical photographs of surfactant-free water/toluene emulsion (W/O) and Pluronic  
 448 F-127 emulsified toluene/water emulsion (O/W) (a) before and (b) after separation.

449  
 450 *3.9. Surface energy estimate of MCFO/RGO/PU sponge*

451 Determining the free-energy of solid surfaces has been an elusive attempt for surface  
 452 scientists. The surface energy can be measured by Fowkes, Owens–Wendt, Zisman, and  
 453 Girifalco–Good–Fowkes (GGF) approaches [53].

454 **I. Fowkes theory**

455 The Fowkes equation is expressed as follows:

456 
$$\gamma_L(1 + \cos \theta) = 2(\gamma_S^d \cdot \gamma_L^d)^{\frac{1}{2}} + 2(\gamma_S^p \cdot \gamma_L^p)^{\frac{1}{2}} \quad (2)$$

457 where  $\theta$  is the contact angle of a liquid droplet on the solid surface,  $\gamma_L$  and  $\gamma_S$  refer  
 458 respectively to surface energies of the liquid and solid calculated using Eq.3,  $d$  and  $h$   
 459 superscripts correspond to the dispersive and polar forces, respectively.

$$460 \quad \gamma_{L(S)} = \gamma_{L(S)}^d + \gamma_{L(S)}^p \quad (3)$$

461 In order to calculate the surface energy, polar and nonpolar liquids should be considered. The  
 462 values of  $\gamma_L^d$  and  $\gamma_L^p$  can be obtained from the literature [54]. The surface energy of the  
 463 magnetic MCFO/RGO/PU sponge was determined by testing different combinations  
 464 including water/diiodomethane, water/hexadecane, glycerol/diiodomethane, and  
 465 glycerol/hexadecane. The average contact angle values of 4  $\mu$ L liquid droplets are reported in  
 466 **Table 1**. Upon substitution of these values into Eq. 2, two equations with two unknowns of  
 467  $\gamma_S^p$  and  $\gamma_S^d$  were solved. The surface energy,  $\gamma_S$ , of the magnetic MCFO/RGO/PU sponge was  
 468 calculated as 48.36 mN·m<sup>-1</sup> ( $\gamma_S^p = 9.16$  mN·m<sup>-1</sup> and  $\gamma_S^d = 39.2$  mN·m<sup>-1</sup>). The surface tension  
 469 values of the polar and nonpolar liquids are exhibited in **Table 1**.

470 **Table 1.** Surface tension of various liquids and their corresponding contact angles (CAs) on  
 471 the magnetic MCFO/RGO/PU sponge surface.

Liquid	$\gamma_L^p$ /m N·m <sup>-1</sup>	$\gamma_L^d$ /mN·m <sup>-1</sup>	$\gamma_L$ /mN·m <sup>-1</sup>	CA ( $\pm I^\circ$ )
Water	51.0	21.8	72.8	165
Glycerol	26.4	37.0	63.4	92
DI Water +Ethanol (95/05%)	-	-	59.0	85
Formamide	18.7	39.5	58.2	84
DI Water +Ethanol (90/10%)	-	-	52.0	75
Ethylene-glycol	19.0	29.0	48.0	70
DI Water +Ethanol (85/15%)	-	-	46.0	66
DI Water +Ethanol (80/20%)	-	-	43.0	60
Ethanol	2.6	18.8	22.4	0
Diiodomethane	0.0	50.8	50.8	0
Hexadecane	0.0	27.6	27.6	0

Dodecane	0.0	25.4	25.4	0
Decan	0.0	23.8	23.8	0

## II. Owens–Wendt theory

The Owens–Wendt theory or extended Fowkes equation:

$$\frac{\gamma_L(1 + \cos \theta)}{2 (\gamma_L^d)^{\frac{1}{2}}} = (\gamma_S^p)^{\frac{1}{2}} \frac{(\gamma_L^p)^{\frac{1}{2}}}{(\gamma_L^d)^{\frac{1}{2}}} + (\gamma_S^d)^{\frac{1}{2}} \quad (4)$$

The surface energy of the magnetic MCFO/RGO/PU sponge was determined by testing different combinations including water/diiodomethane, water/hexadecane, glycerol/diiodomethane, and glycerol/hexadecane. The surface energy,  $\gamma_S$ , of the magnetic MCFO/RGO/PU sponge was calculated as  $49.4 \text{ mN}\cdot\text{m}^{-1}$  ( $\gamma_S^p = 9.7 \text{ mN}\cdot\text{m}^{-1}$  and  $\gamma_S^d = 39.7 \text{ mN}\cdot\text{m}^{-1}$ ).

## III. Zisman model

The Zisman equation is expressed as follows:

$$\gamma_s = \frac{(b\gamma_c + 1)^2}{4b} \quad (5)$$

where,  $b$  is the slope of the fitted line.

**Figure S12** displays the linear fitting of data by the Zisman equation of  $\cos \theta$  versus  $\gamma_L$  for liquids including glycerol, DI water + ethanol (95/05%), ethylene-glycol, DI water + ethanol (90/10%), DI water + ethanol (85/15%), DI water + ethanol (80/20%), and ethanol. The values of critical surface tension of wetting,  $\gamma_c$ , was obtained by extrapolating the linear curve to  $\cos \theta = 1$ . Finally, the surface energy,  $\gamma_S$ , of the magnetic MCFO/RGO/PU sponge was calculated using the Zisman equation ( $23.5 \text{ mN}\cdot\text{m}^{-1}$ ).

## IV. GGF theory

The GGF equation is expressed as follows:

$$\cos \theta = -1 + 2\left(\frac{\gamma_S^d}{\gamma_L}\right)^{\frac{1}{2}} \quad (6)$$

The  $\gamma_S^d$  of the magnetic MCFO/RGO/PU sponge was determined by testing alkane liquids including decane, dodecane, and hexadecane. The  $\gamma_S^d$  of the magnetic MCFO/RGO/PU sponge was calculated as 23.2 mN·m<sup>-1</sup>.

These results clearly revealed that the Fowkes, Owens–Wendt, Zisman, and GGF methods are significantly dependent upon the choice of liquids for calculating the surface energy of a solid surface. In addition, Zisman method works well for nonpolar surfaces. Therefore, since the PU surface has some polar components, the surface energy calculated using the Zisman method is not accurate. Besides, the GGF and Zisman methods underestimate the dispersive component of the magnetic sponge.

The results indicate that the surface energy values, obtained from the Owens–Wendt and the Fowkes methods, are more accurate. In addition, it is obvious that the degree of wettability was mainly dependent on the surface tension of the liquids. Indeed, the contact angle of liquids increases and follows a linear relation with increasing the surface tension. As a result, the magnetic MCFO/RGO/PU sponge could extensively adsorb liquids with low surface tension and repel liquids especially water with high surface tension. On the other hand, although the degree of wettability was mainly dependent on the surface tension of the liquids, there are other factors that affect the WCA such as the liquid density. There is an obvious difference in density of ethylene glycol (1.11 g/cm<sup>3</sup>) and diiodomethane (3.32 g/cm<sup>3</sup>). Therefore, WCA does not follow a linear relationship with surface tension for these two liquids. The contact angle of ethylene glycol is 70° while that of diiodomethane is about 0°.

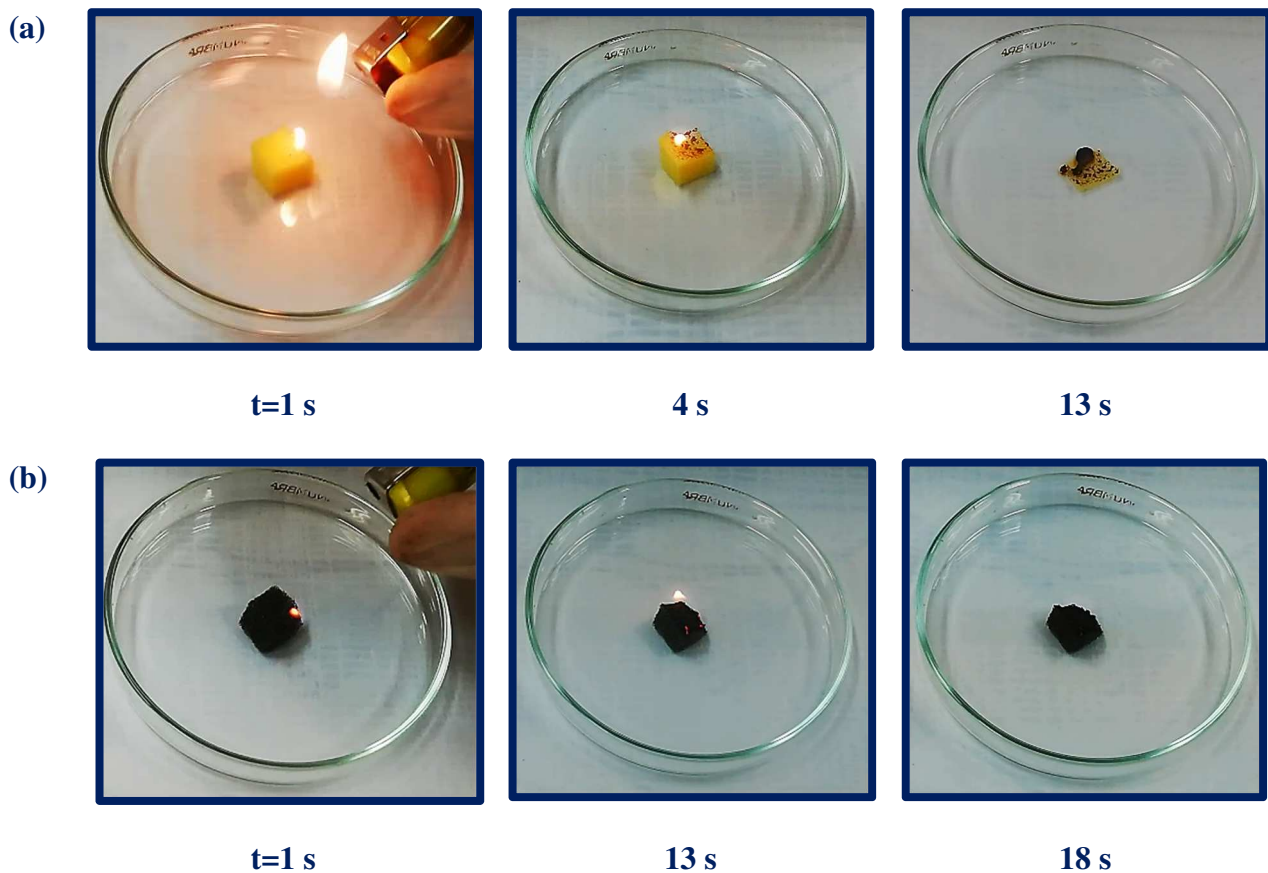
### 3.10. Flame retardancy

PU sponges generate a large extent of toxic gases like carbon monoxide and hydrogen cyanide. Their open-cell structure makes oxygen diffusing through the pores and provides

517 more combustion sites. Hence, it is challenging to produce a highly efficient flame retardant  
518 (FR) barrier with sufficient thickness to strengthen PU sponges against flame [55]. Generally,  
519 during the combustion of FRs, a charred layer is produced which can serve as a physical  
520 barrier to hinder heat from underlayer substrates. Recently, graphene and its derivatives like  
521 RGO have attracted much attention as FR to hinder the transfer of combustion gases into the  
522 inflammable polymer skeleton [15]. Indeed, because of the large surface area and lamellar  
523 structure of RGO, it is favorable to form a physical protective layer on the PU surface.  
524 However, its low decomposition temperature (less than 600 °C) greatly restricts RGO  
525 application as a fire retardant [56]. To address this issue, metal ions and metal oxides known  
526 as non-intumescent flame-retardant materials are used to reinforce the fire retardancy property  
527 of RGO. In addition, they can increase the compatibility of RGO with the PU skeleton [57]. In  
528 this work, the neat PU and MCFO/RGO/PU sponges were cut into equal size, exposed to fire  
529 and the combustion process was recorded. As shown in **Figure 8a**, the neat PU was ignited  
530 immediately and burned completely within 13 s by leaving a small amount of residue at the  
531 end of the test. Compared with neat PU sponge, the modified MCFO/RGO/PU sponge  
532 exhibited self-extinguished and greater thermal stability properties brought by RGO and  
533  $Mn_{0.01}Co_{0.90}Fe_3O_4$  content (**Figure 8b** and **Movie S5**).

534 **Figure S13** illustrates the FTIR spectra of MCFO/RGO/PU sponge before and after  
535 burning. It is obvious that there is no significant change in the FTIR spectrum of  
536 MCFO/RGO/PU sponge after burning test. Therefore, benefiting from the synergistic fire  
537 retardancy of MCFO/RGO nanocomposite and PU sponge, the modified sponge revealed  
538 remarkable adaptability in fire environments.

539



**Figure 8.** The snapshots of flammability test: (a) neat PU, (b) MCFO/RGO/PU sponges.

### 3.11. Mechanical stability of MCFO/RGO/PU sponge

The mechanical properties of MCFO/RGO/PU sponge are very important for practical applications in oil sorption. As shown in **Figure S14a-c**, water droplet still maintained a spherical shape under different mechanical deformation, including multiple times of substrate bending, compressing or twisting. The abrasion resistance of the MCFO/RGO/PU sponge (3×3×2 cm) was also examined. In this test, 200 g load was applied and the sponge was moved back/forth (an abrasion cycle) with abrasion length of about 5 cm. The WCA of the MCFO/RGO/PU sponge was recorded after various abrasion cycles. As shown in **Figure S15**, the WCA did not change obviously after every five abrasion cycles for a total of 20 cycles, indicating that the modified sponge possessed excellent superhydrophobicity and abrasion resistance. The insets in **Figure S15** show that the water droplets on MCFO/RGO/PU sponge still preserved a spherical shape with a WCA of 165°. This result confirmed that the

554 MCFO/RGO/PU sponge maintained its robust superhydrophobicity, because of the strong  
555 adhesion of the MCFO/RGO nanocomposite on the PU sponge surface.

#### 556 **4. Conclusion**

557 A magnetic, superhydrophobic/superoleophilic, and flame-retardant  $Mn_{0.01}Co_{0.90}$   
558  $Fe_3O_4$ /reduced graphene oxide/polyurethane (MCFO/RGO/PU) sponge was fabricated  
559 through a simple dip-coating method.  $Mn_{0.01}Co_{0.90}$   $Fe_3O_4$ /reduced graphene oxide  
560 (MCFO/RGO) nanocomposite was synthesized through *in situ* growth of magnetic MCFO  
561 nanoparticles on RGO *via* a solvothermal method. The nanocomposite decorated on PU  
562 skeleton could increase effectively the surface roughness and change the sponge wettability.  
563 The MCFO/RGO/PU sponge could be easily manipulated by a magnet and separate oil  
564 droplets from surfactant-free and stabilized toluene/water emulsions. Also, it could adsorb oils  
565 and organic solvents up to 131 times of its weight compared to neat PU sponge, while  
566 retaining its ability even after 10 sorption cycles. Additionally, the MCFO/RGO/PU sponge  
567 exhibited good mechanical and abrasion stability. Compared with neat PU, the robust  
568 magnetic MCFO/RGO/PU sponge exhibits self-extinguished property due to the presence of  
569 RGO in the nanocomposite anchored on PU matrix. The reported features make the modified  
570 PU as a promising candidate for remediation of oil spills.

#### 571 **Declaration of Competing Interest**

572 The authors declare that they have no known competing financial interests or personal  
573 relationships that could have appeared to influence the work reported in this paper.

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580 the LASIRE laboratory.

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