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Atmospheric nitrous acid measurement in the French Landes forest

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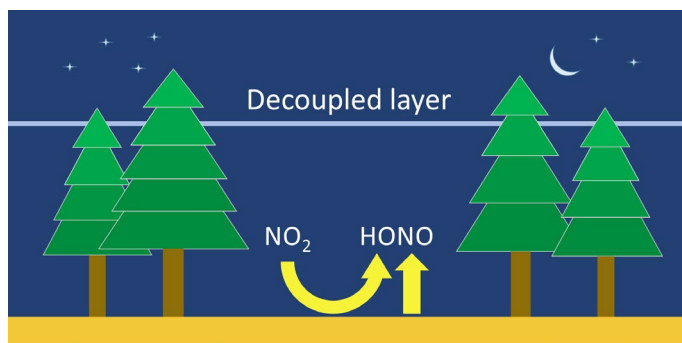
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TOC Figure



16 **ABSTRACT**

17 Nitrous acid (HONO) is a major hydroxyl radical (OH) source, which plays a central role in
18 atmospheric chemistry. In this study, continuous measurements of HONO and other related
19 parameters were conducted from 5th to 25th July 2018 in the French Landes forest. The mean
20 daytime and nighttime HONO concentrations were 60 ppt and 139 ppt, respectively, whilst the
21 mean daily OH production rate from HONO photolysis below the forest canopy was estimated to
22 be 7.6×10^4 molecules $\text{cm}^{-3} \text{s}^{-1}$, much less than those from ozone (O_3) photolysis in the presence
23 of water vapor (3.1×10^5 molecules $\text{cm}^{-3} \text{s}^{-1}$) and alkene ozonolysis (4.4×10^6 molecules $\text{cm}^{-3} \text{s}^{-1}$).
24 Interestingly, the HONO levels were daily peaking between 18:00 to 24:00, that cannot be
25 explained by the known chemical formation pathways. Analysis of collocated measurements
26 suggested that the accumulation of HONO during the evening might be explained by
27 micrometeorological processes. Hence creation of a stable decoupled layer retaining the HONO
28 emission from the soil was associated with the formation of a convective decoupled layer.
29 Increasing relative humidity within this period enhanced HONO gaseous deposition leading to its
30 fast depletion.

31 **KEYWORDS**

32 nitrous acid, Landes forest, atmospheric photochemistry, micrometeorological process, decoupled
33 layer, soil emission

34

35 **Introduction**

36 Nitrous acid (HONO) plays an important role in atmospheric photochemistry due to its fast
37 photolysis yielding hydroxyl radicals (OH) and nitric oxide (NO), both of which contribute to the
38 ground-level ozone (O₃) and aerosol production.¹⁻⁴ Many field and modeling studies have shown
39 that HONO photolysis can contribute to 25-56% of OH radical production in the daytime,⁵⁻⁷ even
40 reaching up to 80% in the early morning.⁸ Despite the significance of HONO to atmospheric
41 chemistry, a molecular-level understanding of its sources remains unclear.^{6,9}

42 During the past decades, a large number of ambient HONO measurements have been
43 conducted at various sites including rural, suburban, urban, forest, mountain, ocean and even in
44 polar regions, showing that HONO concentration ranges from several ppt to more than 10 ppb.^{6,10-}
45 ¹⁸ Anthropogenic direct emissions such as vehicle exhaust with emission ratios of HONO/NO_x
46 ranging from 0.3% to 1.6% and natural sources including soil emission are regarded as important
47 HONO sources for urban/suburban, and rural/forest regions, respectively.¹⁹⁻²¹ The study of Su et
48 al.¹⁹ showed that HONO can be released from nitrite degradation present in the soil and might
49 strongly increase globally due to the enhanced fertilizer use and soil acidification. However, soil
50 biological activity has been recently suggested as the most important HONO source from soil
51 emission.²² This process can release HONO in greater quantities than from the acid-base and
52 Henry's law equilibria of the aqueous phase in soil, but its production is influenced by the soil
53 temperature and water content as a result of their influence on microbial activity.

54 In addition to primary emissions, the atmospheric secondary reactions including gas phase
55 reaction of NO with OH radicals^{9,23} and heterogeneous reactions of NO₂ on various surfaces such
56 as ground, leaf, etc., are also suggested to significantly contribute to HONO formation.² In recent
57 years, many laboratory studies have shown that photolysis of nitric acid (HNO₃) and nitrate ions

58 (NO_3^-) can be an important daytime source of HONO.²⁴⁻²⁶ In addition, HONO production from
59 aqueous nitrate photolysis can be enhanced by dissolved aliphatic organic matter due to the reaction
60 of O_2^- with NO_x producing nitrite (NO_2^-).²⁷ Moreover, the photosensitized conversion of NO_2 on
61 surfaces was also proved to be an efficient HONO formation pathway,^{28,29} whilst the ozonolysis
62 and photolysis of organic nitrogen substances, such as neonicotinoid nitenpyram ozonolysis³⁰
63 and ortho-nitrophenols³¹ and nitrotyrosine photolysis,³² are also suggested to be responsible for the
64 formation of HONO. These findings are helpful to fill the gaps between the field observations and
65 numerical models. However, the sources and fate of HONO in ambient conditions remain still
66 unclear.⁹

67 In addition to these direct emission and formation pathways, the local micrometeorological
68 processes can also significantly influence the primary pollutant concentrations and should not be
69 neglected. As an example, it has been recently reported that the formation of a shallow decoupled
70 layer below the canopy, can lead to a significant increase of monoterpenes and monoterpene
71 oxidation products near ground level in the Finnish boreal forest³³ and the French Landes forest.³⁴

72 This work aims to contribute to unravel biological, chemical, physical and
73 micrometeorological processes on HONO sources and sinks in the context of a forested site
74 characterized by high levels of biogenic volatile organic compounds (BVOCs) and weakly
75 influenced by anthropogenic emissions (i.e., low NO_x air masses). Measurements of atmospheric
76 HONO as well as meteorological parameters, trace gases, and volatile organic compounds (VOCs)
77 were performed in a maritime pine forest located in the French Landes (southwestern part of
78 France), previously reported as a site highlighting high monoterpene concentrations.^{35,36} Very
79 surprisingly, a daily high HONO peak ranging from 18:00 to midnight was observed. The
80 contribution of HONO photolysis to the primary OH production was evaluated and compared to

81 those of O₃ photolysis and reaction of the O(¹D) product with water vapor as well as alkene
82 ozonolysis. Finally, a HONO budget including soil emission, chemical formation and loss
83 pathways as well as local micrometeorological processes was estimated and analyzed for the
84 summer of 2018.

85 **Experimental section**

86 **Site description.** The measurements presented in this study were performed as part of the
87 CERVOLAND (Characterization of Emissions and Reactivity of Volatile Organic Compounds in
88 the Landes forest) campaign and conducted from 5th to 25th July 2018. The measurement site was
89 located next to the European Integrated Carbon Observation System (ICOS) station at Bilos in the
90 Landes forest (44°29'39.69" N, 0°57'21.75" W), which is around 40 km far from the nearest urban
91 area of the Bordeaux metropole and 25 km from the North Atlantic Oceanic. More descriptions
92 about the site were given in previous studies.^{35,37,38} Briefly, this field site is mainly covered by
93 maritime pine trees (*Pinus pinaster* Aiton) and weakly influenced by anthropogenic pollution. It is
94 characterized by very high levels of biogenic VOCs (up to 20 ppb for the total monoterpene
95 concentration)^{39,36} and the soil is a sandy acidic hydromorphic podzol with a discontinuous layer
96 of iron hard pan at a 75 cm depth.³⁴

97 **Measurement techniques.** Atmospheric HONO was measured by a commercial long path
98 absorption photometer instrument (QUMA, Model LOPAP-03). A detailed description of the
99 LOPAP instrument can be found elsewhere.^{27,40} In brief, the sampling stripping coil (at 4 m above
100 the ground level (a.g.l.)) was kept at a constant temperature of 15°C and consists of two channels:
101 channel 1 uptakes HONO and a small fraction of interfering species, while the interfering species
102 are quantified by channel 2.⁴¹ The ambient air sampling flow rate and peristaltic pump velocity

103 were set to 1.1 L min^{-1} and 10 revolutions per minute, with a HONO collection efficiency of 99.9%
104 and an uncertainty of 10%. The LOPAP was calibrated by a known concentration nitrite (NO_2^- ,
105 Titrisol, $1000 \text{ mg L}^{-1} \text{ NO}_2^-$ in water) solution every 4 days, whilst zero measurements were
106 conducted every 8 hours by injecting pure N_2 . Detection limit of the LOPAP was estimated to be
107 0.3 ppt (i.e., 2 times noise of zero measurements). Then, HONO concentration was obtained by
108 subtracting the calibrated signal of channel 2 from channel 1 with a time resolution of 30 s.

109 In addition, the Vocus proton-transfer-reaction mass spectrometry (PTR-MS), with a high
110 time response and sensitivity (detection limit $< 1 \text{ ppt}$), was used to measure VOCs during this field
111 campaign. More descriptions about the setup and the results of Vocus PTR-MS were reported
112 elsewhere.^{36,42} Additionally, the photolysis frequencies of NO_2 and O_3 represented by J_{NO_2} and J_{O_3}
113 were quantified by a spectroradiometer (Meteorologie Consult GmbH 6007) in a clearing. The
114 photolysis frequency of HONO (J_{HONO}) was obtained by multiplying J_{NO_2} by the ratio of calculated
115 J_{HONO} and J_{NO_2} based on the NCAR Tropospheric Ultraviolet and Visible (TUV) model.^{43,44} NO_x
116 and ozone (O_3) concentrations were measured at 4 m a.g.l. by chemiluminescence and UV
117 absorption analyzers, respectively, both of which have a detection limit of 0.5 ppb . The
118 meteorological parameters such as temperature (T), relative humidity (RH), wind speed (WS),
119 wind direction as well as CO_2 concentrations at 0.46, 1.6, 3.4, 5.7, 8.5, 11.8, and 15.6 m a.g.l. were
120 continuously monitored, while the solar radiation was measured at 15.6 m a.g.l. and measurement
121 data were provided by the European Integrated Carbon Observation System (ICOS) station. All
122 data are reported and discussed in Coordinated Universal Time (UTC) in this study, two hours later
123 than local time.

124 **Parametrization of OH concentration.** The steady OH concentration ($[\text{OH}]_{\text{ss}}$) could be estimated
125 using Eq. 1, as in the study by Mermet et al.:³⁷

126
$$[OH]_{ss} = 1 \times 10^5 + \frac{Rad_t \times K_{a,t}}{1 \times 10^3} \times [[OH]_{max} - 1 \times 10^5] \quad (\text{Eq. 1})$$

127 Notably, 1×10^5 molecules cm^{-3} was retained as a typical value for OH concentration in forest
 128 environments at nighttime.⁴⁵⁻⁴⁹ The Rad_t (W m^{-2}) was the solar irradiation above the forest canopy
 129 and $K_{a,t}$ was the ratio between the solar irradiance values measured below and above the canopy as
 130 previously proposed by Mermet et al.,³⁷ whilst $1 \times 10^3 \text{ W m}^{-2}$ was the maximum of solar irradiation
 131 during the field campaign. Besides, 6×10^6 molecules cm^{-3} measured by Birmili et al.⁴⁸ and $3 \times$
 132 10^6 molecules cm^{-3} reported by Hens et al.⁴⁵ were used as the high and lower limits of $[OH]_{max}$,
 133 respectively.³⁷

134 In addition, another empirical formula proposed by Ehhalt & Rohrer (2000)⁵¹ was widely
 135 used to estimate the $[OH]_{ss}$ concentration in the daytime:^{15,44,52}

136
$$[OH]_{ss} = a(J_{O1D})^\alpha (J_{NO2})^\beta \frac{bNO_2+1}{cNO_2^2+dNO_2+1} \quad (\text{Eq. 2})$$

137
$$(\alpha = 0.83, \beta = 0.19, a = 4.1 \times 10^9, b = 140, c = 0.41, \text{ and } d = 1.7)$$

138 If alkene ozonolysis dominates the production of OH radicals, notably during nighttime, $[OH]_{ss}$
 139 concentration can be calculated using the following equation:^{36,53}

140
$$[OH]_{ss} = \frac{\sum k_{(O_3+VOC_i)} \alpha_i [O_3] [alkene_i]}{\sum k_{(OH+VOC_i)} [alkene_i] + k_{(OH+O_3)} [O_3]} \quad (\text{Eq. 3})$$

141 where $k_{(O_3+VOC)}$ is the rate constant for the reaction of O_3 with alkenes with an OH yield of α ;
 142 $k_{(OH+VOC)}$ and $k_{(OH+O_3)}$ are the rate constants for OH reacting with alkenes and O_3 , respectively. The
 143 main alkenes are monoterpenes in the Landes forest region, but also consist of a small amount of
 144 isoprene and sesquiterpenes (Figures S1-S2).³⁶ In addition, a previous field observation in 2017
 145 summer on the same site showed that the identified species including α -pinene, β -pinene, camphene,
 146 sabinene, myrcene, 2-carene, α -phellandrene, Δ^3 -carene, limonene, ocimene, α -terpinene,
 147 terpinolene account for $\sim 80\%$ of the total monoterpene concentrations, whilst β -caryophyllene

148 represents ~40% of the total sesquiterpenes and ~90% of total sesquiterpene reactivity.³⁷ Therefore,
149 these identified species including their accounted percentages were selected for the calculations Eq.
150 3, whilst their reaction rate constants with OH radicals and O₃ as well as their corresponding OH
151 radical production yields were taken from previous studies.⁵⁴⁻⁵⁶ Notably, OH production produced
152 by the reactions of hydroperoxyl radical (HO₂) with NO or O₃ was neglected, since NO
153 concentration was below the detection limit of NO_x analyzer most of the time, and HO₂
154 concentration was not measured in this study.

155 **RESULTS AND DISCUSSION**

156 **Measurements of meteorological conditions and trace gases.** Figure 1 displays an overview of
157 meteorological conditions (measured at 3.4 m a.g.l. except for solar radiation measured at 15.6 m
158 a.g.l.) and the measured HONO, NO_x, and O₃ concentrations, for which diel variation pattern was
159 consistent during the campaign.³⁵ Except for some rainy and cloudy days (i.e., 6th, 21st, and 22nd
160 July), the weather remained sunny with strong solar irradiation, indicating favorable conditions for
161 intense photochemical activities. The air masses within the forest canopy were relatively stable,
162 with wind speed < 1 m s⁻¹, suggesting a major influence of local emissions. The ambient
163 temperature ranged from 11 °C to 35 °C with a mean value (± standard deviation) of 22 ± 6 °C and
164 an average RH of 74 ± 21 %. Figure 1C shows that the average concentration of NO₂ was 1.6 ± 0.9
165 ppb, while the concentration of NO was below the detection limit (i.e., < 0.5 ppb) most of the time,
166 but showing a peak in the morning, with a maximum value of 5.6 ppb at 5:00 on 10th July. O₃
167 concentration appeared with a mean value of 28 ± 19 ppb, with an average maximum of 49 ppb at
168 12:00 (Figure 2). The HONO concentration ranged from 4 ppt to 1.1 ppb, with a mean value of 96
169 ± 107 ppt, which is slightly higher than typical values of 30 to 90 ppt previously observed in forest
170 environments.^{10,17,57,58}

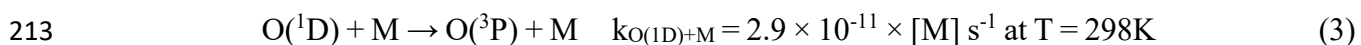
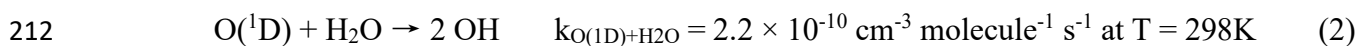
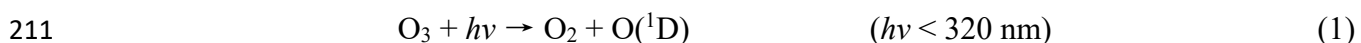
171 The diel profiles of meteorological conditions and trace gases are depicted in Figure 2, all
172 of which showing clear diurnal variations. The CO₂ concentrations measured at different heights
173 from 0.46 m a.g.l. to 15.6 m a.g.l. was used as a proxy to estimate the decoupled layer height
174 (DLH).³³ The temperature, O₃ concentration and photolysis frequencies (i.e., J_{O₃}, J_{NO₂}, and J_{HONO})
175 followed the pattern of the solar irradiation, while the RH and NO₂ concentrations exhibited
176 opposite diel profiles. The mean NO₂ concentration started to increase from 0.9 ± 0.14 ppb at 14:00
177 to 2.3 ± 1.3 ppb at 21:00, and stayed constant during the night. Consistent with the increase of NO₂
178 photolysis rate (Figure 2F) and the expansion of boundary layer height in the early morning, the
179 concentration of NO₂ quickly declined. The average HONO concentrations in the daytime (UTC
180 time, 7:00-19:00) and nighttime (UTC time, 20:00-next day 6:00) were 60 ± 32 ppt and 139 ± 72
181 ppt, respectively. HONO measurements in the Finnish boreal forest also showed higher HONO
182 concentrations at night,⁵⁷ while several other forest measurements exhibited the opposed trend due
183 to unknown daytime HONO sources.^{10,17,59} Interestingly, the mean HONO concentration quickly
184 rose in the evening from 53 ± 29 ppt at 18:00 to 300 ± 209 ppt at 20:00 (Figure 2D), and the fastest
185 increasing case was on 14th July, where the HONO concentration increased from 94 ppt to 1.1 ppb
186 within 2 hours (Figure 1D). It is worth pointing out that such a fast rise has not yet been reported
187 from previous HONO measurements.^{10,17,57,59}

188 **Impact of HONO and O₃ photolysis and alkene ozonolysis on the primary OH production.**

189 Firstly, as shown in Figure 3A, the low and high [OH]_{ss} concentrations calculated by Eq.1 are in
190 the ranges of 1-10 × 10⁵ molecules cm⁻³ and 1-20 × 10⁵ molecules cm⁻³, with the average values of
191 2.9 × 10⁵ molecules cm⁻³ and 4.8 × 10⁵ molecules cm⁻³, respectively. The maximum concentration
192 of [OH]_{ss} is shown at 11:00 owing to the strongest solar irradiation within the forest canopy,
193 consistent with the maximum J_{NO₂}, J_{HONO} and J_{O₃} (Figures 2E and 2F). Notably, all photolysis

194 frequencies below the forest canopy were calculated by the measured ones multiplying $K_{a,t}$ (i.e.,
 195 the ratio between the solar irradiance values measured below and above the canopy). Similarly, the
 196 $[\text{OH}]_{\text{ss}}$ expressed by the NO_2 concentrations and photolysis frequencies of O_3 and NO_2 (i.e.,
 197 calculated by Eq. 2) also showed the maxima at 11:00, with a concentration of 2.4×10^6 molecules
 198 cm^{-3} . It is worth noting that solar irradiation under the forest canopy was weak (Figure 2F) in the
 199 time period of 17:00 to the next day 9:00, both of which suggesting that alkene ozonolysis might
 200 have been the main source of OH radicals. Therefore, $[\text{OH}]_{\text{ss}}$ is calculated by means of Eq. 3,
 201 considering the competition between OH production and removal processes leading to the $[\text{OH}]_{\text{ss}}$
 202 formation. Different to the result calculated by Eq. 1, a greater $[\text{OH}]_{\text{ss}}$ concentration in the range
 203 of $4\text{-}8 \times 10^5$ molecules cm^{-3} was observed during the time period from 17:00 to 24:00, that results
 204 from the subsequent monoterpene ozonolysis. As a result, the average $[\text{OH}]_{\text{ss}}$ concentration ($7.4 \times$
 205 10^5 molecule cm^{-3}) calculated by Eqs. 2-3 is higher than those (low average value: 2.9×10^5
 206 molecule cm^{-3} , high average value: 4.8×10^5 molecule cm^{-3}) by Eq. 1.

207 It has been widely accepted that HONO photolysis is an important source of OH radicals
 208 in the atmosphere, especially in urban and suburban regions.^{5,52,60} In addition, O_3 photolysis in the
 209 presence of water vapor is the most important source of OH radicals globally (reactions 1-3) during
 210 daytime.⁶¹



214 where M is N_2 or O_2 .

215 Then we compared the OH production rates from the photolysis of HONO and O_3 , and alkene
 216 ozonolysis. While considered here as the main OH production pathways yet some others such as

217 aldehyde photolysis or RC(O)O₂ with HO₂ reaction may also contribute.⁶² The OH production rates
 218 from HONO photolysis (P_{OH}(HONO)), O₃ photolysis in the presence of water vapor (P_{OH}(O¹D)),
 219 and alkene ozonolysis (P_{OH}(alkene ozonolysis)) were calculated using equations 4-6,
 220 respectively:^{52,54,61}

$$221 \quad P_{\text{OH}}(\text{HONO}) = J_{\text{HONO}}[\text{HONO}] \quad (\text{Eq. 4})$$

$$222 \quad P_{\text{OH}}(\text{O}^1\text{D}) = \frac{2J(\text{O}^1\text{D})[\text{O}_3]k_{\text{O}^1\text{D}+\text{H}_2\text{O}}[\text{H}_2\text{O}]}{k_{\text{O}^1\text{D}+\text{H}_2\text{O}}[\text{H}_2\text{O}] + k_{\text{O}^1\text{D}+\text{M}}[\text{M}]} \quad (\text{Eq. 5})$$

$$223 \quad P_{\text{OH}}(\text{alkene ozonolysis}) = \sum k_{\text{O}_3+\text{alkene}_i} \alpha_i [\text{alkene}_i][\text{O}_3] \quad (\text{Eq. 6})$$

224 As depicted in Figure 3B, P_{OH}(O¹D) showed a maximum value of 1.6 × 10⁶ molecules cm⁻³
 225 s⁻¹ at 11:00 and a daily average value of 3.1 × 10⁵ molecules cm⁻³ s⁻¹. However, P_{OH}(HONO) was
 226 much less, showing a maximum value of 3.7 × 10⁵ molecules cm⁻³ s⁻¹ at 11:00 and a mean
 227 production rate of 7.6 × 10⁴ molecules cm⁻³ s⁻¹, respectively. On the contrary, as shown in Figure
 228 3B, alkene ozonolysis contributes significantly to the OH radical formation from 18:00 to 24:00
 229 with a mean production rate of 9.1 × 10⁶ molecules cm⁻³ s⁻¹, therefore, showing a daily average
 230 production rate of 4.4 × 10⁶ molecules cm⁻³ s⁻¹, higher than P_{OH}(O¹D) and P_{OH}(HONO). However,
 231 it should be noted that even if alkene ozonolysis contributed significantly to OH formation, [OH]_{ss}
 232 still stayed at a relatively low level as a result of an active OH chemistry during this period,
 233 confirmed by the formation of alkene-derived oxygenated VOCs.³⁶ All in all, this result shows P_{OH}
 234 (HONO) was insignificant in this area, only accounting for 1.6% of total daily OH production rate
 235 from these different formation pathways, probably due to the low concentration of HONO in the
 236 daytime and relatively weak solar irradiation below the forest canopy compared to that above the
 237 forest canopy.

238 **HONO budget analysis from 18:00 to 24:00.** As discussed above, surprisingly, a daily high but
239 unexplained HONO peak was observed from 18:00 to 24:00. Therefore, to identify the potential
240 sources, the HONO budget was investigated during this period. Because the photolysis frequencies
241 below the forest canopy were negligible (Figures 2E and 2F), the photochemical sources (i.e.,
242 $\text{NO}_3^-/\text{HNO}_3$ photolysis, photosensitized oxidation of NO_2 , etc.) and photolysis of HONO were not
243 considered. As a result, the main factors affecting the ambient HONO concentration from 18:00 to
244 24:00 were summarized by Eq. 7:

$$245 \quad \frac{d[\text{HONO}]}{dt} = P_{\text{het}} + P_{\text{OH+NO}} + P_{\text{emis}} + P_{\text{unknown}} - L_{\text{dep}} - L_{\text{OH+HONO}} \quad (\text{Eq. 7})$$

246 where, $d[\text{HONO}]/dt$ is the observed change rate of HONO concentration; HONO sources include
247 heterogeneous conversion of NO_2 to HONO (P_{het}), gas-phase reaction of OH and NO ($P_{\text{OH+NO}}$),
248 direct emission (P_{emis}) and unknown sources including some processes influencing HONO
249 concentration (P_{unknown}), while the loss pathways include dry deposition of HONO (L_{dep}) and
250 reaction between OH and HONO ($L_{\text{OH+HONO}}$).

251 Figure 4 shows the average rates of $P_{\text{OH+NO}}$, P_{het} , $L_{\text{OH+HONO}}$, L_{dep} , and the sum of P_{emis} and
252 P_{unknown} during the period of 18:00-24:00, and the details on their calculations were described in
253 the Supplemental Materials Text S1. First, the $L_{\text{OH+HONO}}$ and $P_{\text{OH+NO}}$ rates were negligible
254 compared to other formation and loss pathways. Both P_{het} and L_{dep} rates were influenced by the
255 DLH, which was around 6 m from 18:00 to 24:00, based on the separation of vertical CO_2
256 concentrations (Figure 2A) and temperature inversion (Figure 5). Notably, DLH includes stable
257 decoupled layer height (SDL) and convective decoupled layer height (CDL) in the present study.
258 Although the adopted value of DLH can impact the rates of P_{het} and L_{dep} , their ratio (i.e., $P_{\text{het}}/L_{\text{dep}}$)
259 remained constant. P_{emis} was not available in this study, but the sum of P_{emis} and P_{unknown} were
260 calculated based on the Eq. 7. As shown in Figure 4, the known HONO sources, i.e., gas phase

261 reaction of NO with OH radicals and heterogeneous conversion of NO₂ on surfaces, cannot explain
262 the fast increase of HONO ($0.7\text{-}1.0 \times 10^6$ molecules cm⁻³ s⁻¹) during the nightfall (Figure 2D).

263 Interestingly, there is a temperature inversion layer formed at the ground level (~ 6 m) at ~
264 17:00 (Figure 5), caused likely by radiation cooling.^{63,64} This temperature inversion can yield a
265 stable decoupled layer (SDL) close to the ground and retain the emissions from the ground and the
266 vegetation within this thin layer,^{33,65} concomitantly to the rapid increases of HONO (Figure 2D),
267 NO₂ (Figure 2C) and monoterpene concentrations (Figure S2). However, the NO concentration
268 stayed low from 18:00 to 24:00 due to the rapid consumption by peroxy radicals formed from the
269 terpene ozonolysis and O₃, producing organic nitrates (see Figure S3)³⁶ and NO₂,⁶¹ respectively.
270 However, HONO concentration did not increase significantly at 17:00, due to its remaining strong
271 photolysis (Figure 2E). Oswald et al.²² showed that HONO emissions from soils were greater when
272 soil temperature increases (20-30 °C), but less as soil water content increases (> 10% of water
273 holding capacity). In this study, the gaseous HONO level showed strong positive and negative
274 correlations with the ambient temperature (Figure S4A) and RH (Figure S5A), respectively. This
275 is suggesting that soil emission can be potentially an important HONO source within the Landes
276 forest. For example, the highest HONO peak detected during the field campaign might be related
277 to the relatively higher temperature and lower RH in the evening on July 14th (Figure 1). The
278 exchange of NO_x and HONO between the various surfaces present (such as vegetation cover) and
279 the atmosphere may be bi-directional, which cannot be distinguished in this study.

280 Afterwards, a weak convective decoupled layer (CDL) was created at the ground level after
281 20:00 and remained stable at 22:00 until 6:00 of the next day (Figures 5D and 5E). Notably, the
282 correlation between HONO and temperature became weaker (Figure S4B) but HONO
283 concentration still showed a strong negative correlation with ambient RH (Figure S5B), suggesting

284 that increasing RH likely enhanced the deposition loss of gaseous HONO in addition to its negative
285 influence on HONO emission from soil. Therefore, both the created weak CDL at the ground level
286 and the increasing RH (Figure 1B) probably led to the fast depletion of gaseous HONO from 20:00
287 to 24:00 (Figure 2D). However, the NO₂ concentration stayed constant and NO concentration
288 showed an increasing trend during the night, which were probably due to their deposition rate (i.e.,
289 NO₂, $3-8 \times 10^{-3} \text{ m s}^{-1}$; NO, $1-2 \times 10^{-3} \text{ m s}^{-1}$)⁶⁶ lower than HONO ($2 \times 10^{-2} \text{ m s}^{-1}$), and less active O₃
290 chemistry. Interestingly, many isoprene-, monoterpene-, and sesquiterpene-derived oxidation
291 products also show a similar high peak during this period (Figure S6 and the study of Li et al.
292 (2020)). This was caused by high concentrations of terpenes and O₃ within the SDL at the ground
293 level during the period from 17:00 to 20:00. After 20:00, the created CDL started to enhance the
294 dry deposition of these oxidation products, whilst O₃ concentration decreases fast, both of which
295 caused the fast loss of these oxidation products. During the period from 00:00 to 7:00, the gaseous
296 HONO concentration stayed constant, indicating the sinks and sources of HONO reached an
297 equilibrium.

298 Notably, the slope fitted by the linear regression for HONO/NO₂ ratios against time was
299 generally used as the conversion frequency of NO₂-to-HONO (P_{het}) during the
300 nighttime.^{10,15,16,44,52,60,67,68} However, sometimes the creation of a SDL at the ground level can
301 efficiently promote the concentrations of primary pollutants including HONO after sunset, which
302 will overestimate the heterogeneous conversion rate of NO₂-to-HONO. Therefore, chemical
303 formation pathways, direct emissions and meteorological processes should all be taken into
304 consideration to better understand the variations of atmospheric species during field observations.

305 **CONCLUSION**

306 In the atmosphere, HONO is present at trace levels, typically up to a few ppb, but plays an important
307 role in the tropospheric chemistry due to its photolysis, producing one of the major oxidants (i.e.,
308 OH radicals).¹⁻⁴ In the present study, HONO as well as meteorological conditions and other
309 measured trace gases showed regular variations during the campaign except for rainy days. The
310 daily mean HONO concentration was 96 ± 107 ppt and exhibited a large increase from 19:00 to
311 22:00.

312 Terpene ozonolysis contributed significantly to OH radical formation with an average value
313 of 4.4×10^6 molecules $\text{cm}^{-3} \text{s}^{-1}$, higher than O_3 photolysis in the presence of water vapor and HONO
314 photolysis, due to the active alkene ozonolysis chemistry after sunset. However, the daily OH
315 formation rate from HONO photolysis was only around 7.6×10^4 molecules $\text{cm}^{-3} \text{s}^{-1}$ as a result of
316 low HONO concentrations and relative weak solar irradiation below the forest canopy.
317 Interestingly, HONO peaked daily from 18:00-24:00, up to 1 ppb, which was not yet reported for
318 forested environments and could not be explained by the usually reported chemical formation
319 pathways. The creation of an SDL at the ground level retaining HONO produced from the soil
320 within the thick layer leads to a fast HONO increase. Afterwards, both the created CDL and
321 increasing RH enhanced HONO deposition resulting in its fast loss. Overall, the local
322 meteorological conditions are highlighted in this work to be critical parameters to be taken into
323 account as they significantly influence air pollutants at the ground level.

324

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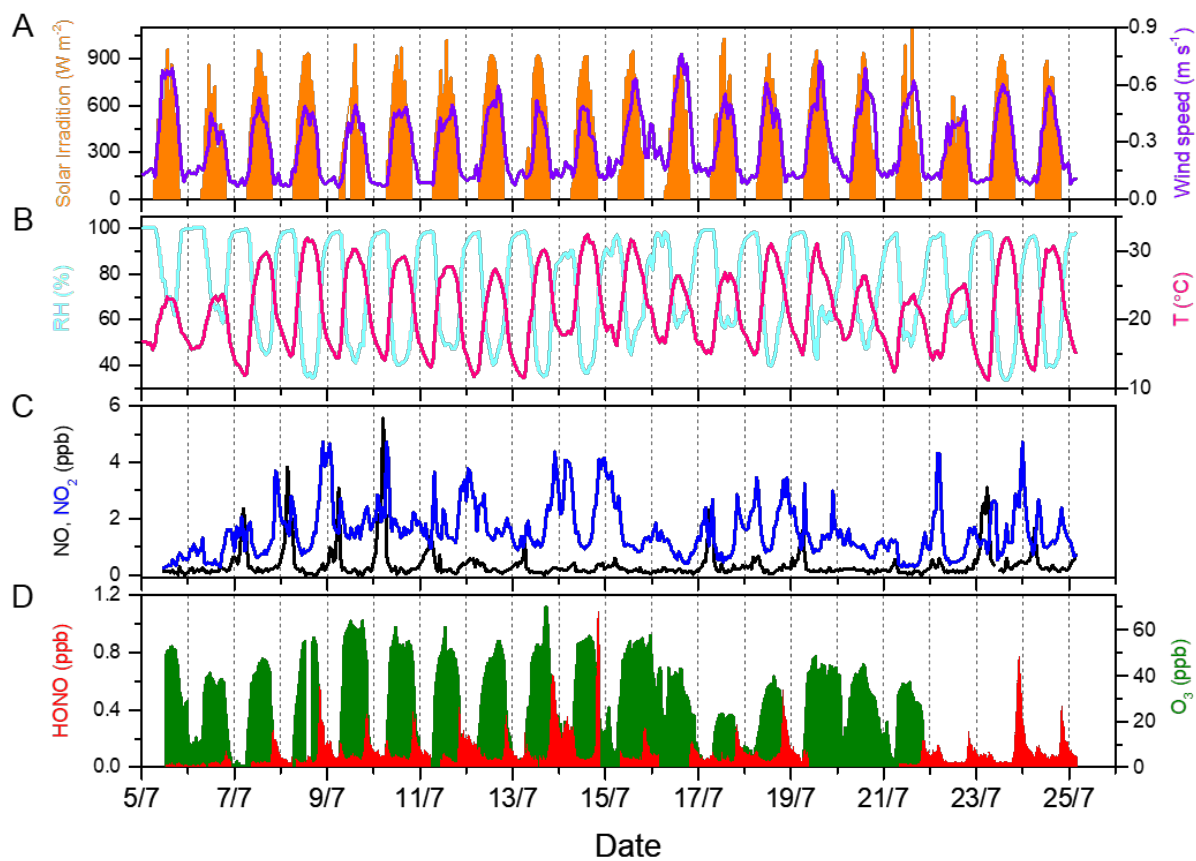
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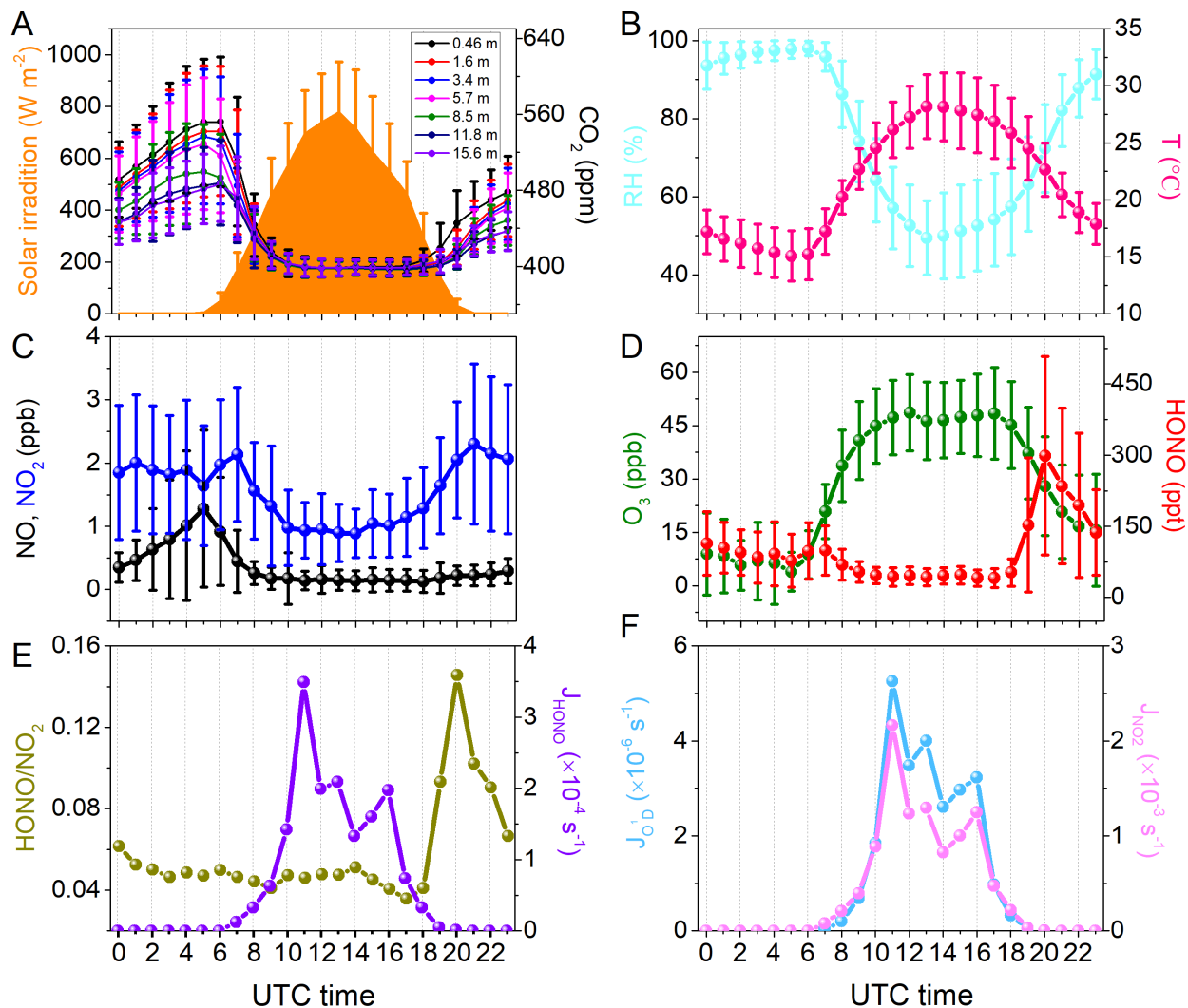
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581 **Figures**



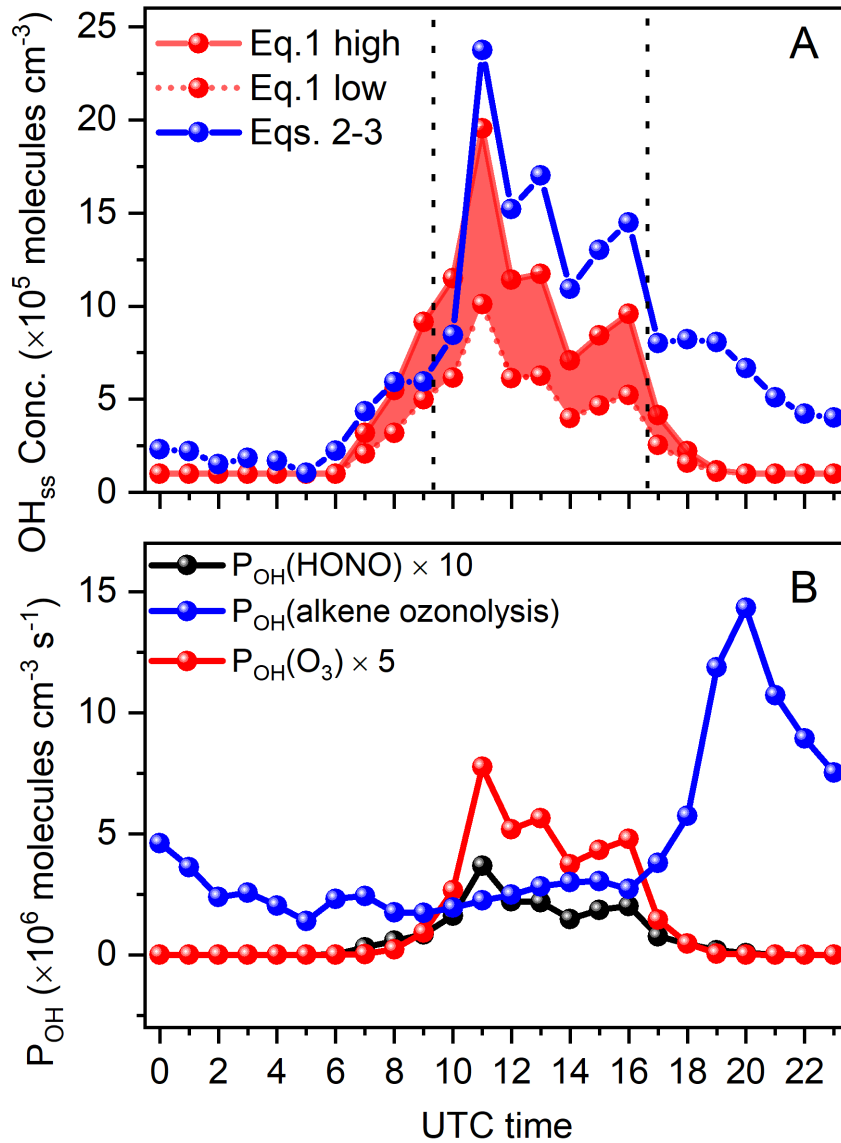
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583 Figure 1. Temporal variations of (A-B) meteorological parameters and measured (C) NO and NO₂,
584 (D) HONO and O₃ during the measurement period from 5th to 25th July, 2018.



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 586 Figure 2. Average diel variations of (A-B) meteorological parameters and CO₂ concentrations at
 587 different heights and measured concentrations of (C) NO_x, (D) HONO and O₃, (E) HONO/NO₂
 588 and calculated J_{HONO}, (F) J_{O1D} and J_{NO2} below the forest canopy during the measurement period.

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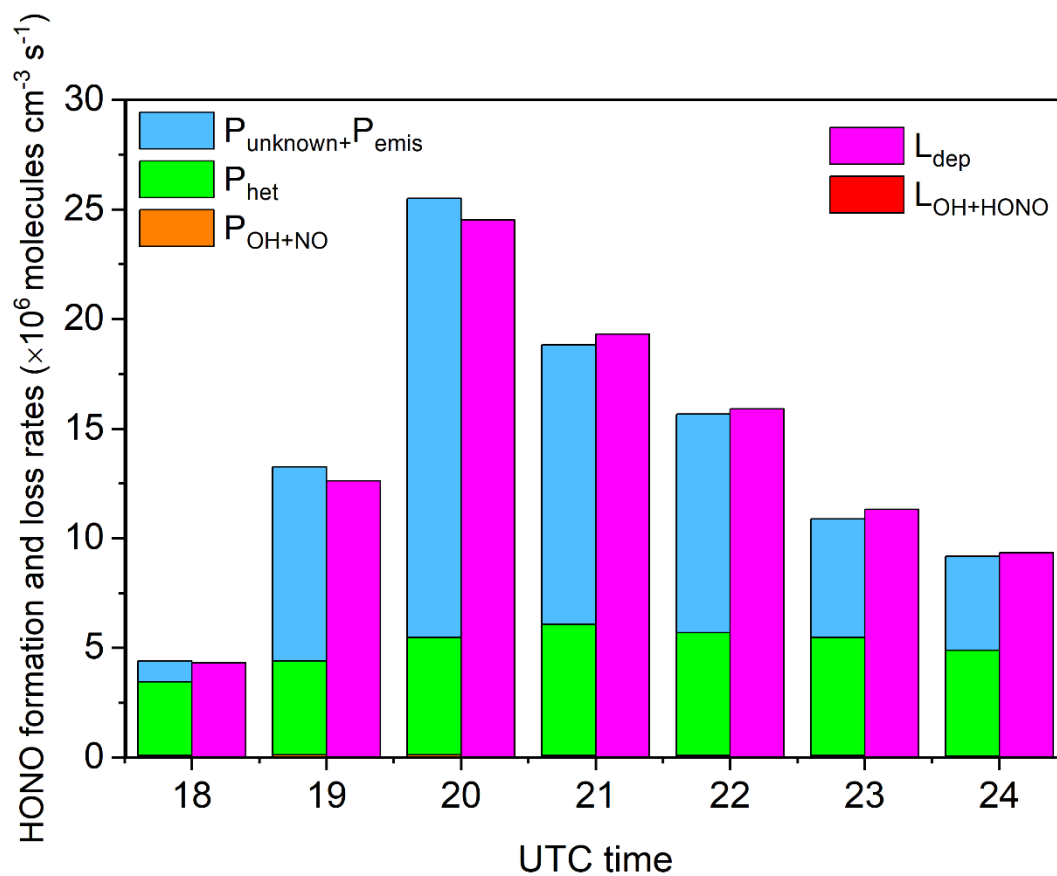


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592 Figure 3. (A) The average steady-state OH concentration and (B) production rates from photolysis

593 of HONO and O_3 , and alkene ozonolysis below the forest canopy during the measurement period.

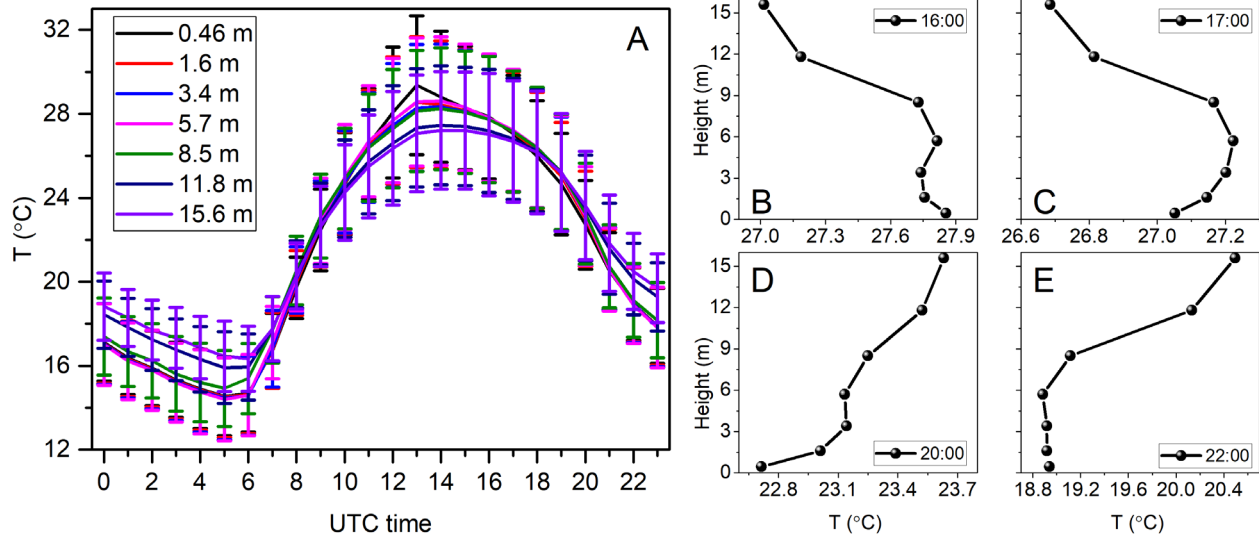
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596 Figure 4. HONO formation and loss rates from different pathways during the measurement period
597 from 5th to 25th July, 2018.

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600 Figure 5. (A) Diurnal variations of vertical temperature and (B-E) the vertical average temperature
 601 at a specific time.