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## A New Type of Spirocyclic Photochromes Reacting with Light of both UV and Visible Ranges†

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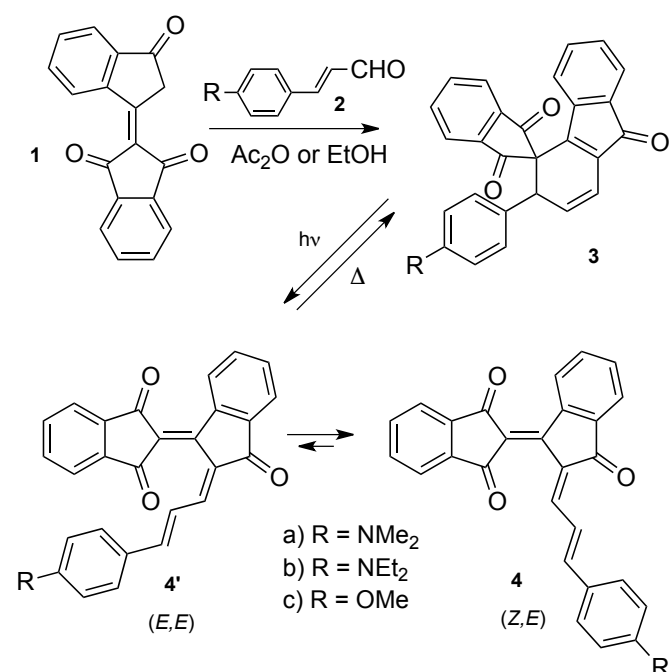
**Condensation of bindone with cinnamic aldehydes produces cyclic isomers instead of the expected colored merocyanine dyes. These compounds exhibit unprecedented photochromic behavior: the formation of the colored open forms is observed upon irradiation by light of any wavelength between 254 and 642 nm.**

During the past decades, merocyanine (MC) dyes<sup>1</sup> are the focus of numerous researches owing to the importance of this type of organic dyes in different fields of materials science, such as nonlinear optics,<sup>2</sup> solar cells,<sup>3</sup> thin-film transistors<sup>4</sup> and biological applications.<sup>5</sup> A special group of merocyanines is related to photochromism. Thus, such important classes of known spiroheterocyclic photochromes<sup>6</sup> as spiropyrans, spirooxazines, chromenes and others form merocyanines as a result of the UV light induced cleavage of the C-O, or more rarely C-C, bonds.<sup>7,8</sup> These merocyanines are usually less stable than the initial cyclic forms and undergo more or less rapid cyclization and the respective color loss. Depending on the structural factors, temperature or media, the equilibrium between the cyclic and merocyanine forms can be shifted toward the latter partially or completely.<sup>9</sup> The stable MC can be isolated and characterized.<sup>10</sup> Some of them can exhibit negative (reverse) photochromism,<sup>11</sup> that is the reversible cyclization upon irradiation by visible light. Recently, the important for the biomedical applications ability of photochromes to react with visible light was demonstrated on the example of substituted azobenzenes.<sup>12</sup>

More than a decade ago, we found that the product of the reaction between bindone (**1**) and *p*-dimethylaminocinnamic aldehyde (**2a**) is not the expected colored merocyanine **4a**, but its cyclic isomer **3a**.<sup>13</sup> In solution or polymer matrix, irradiation

of **3a** at 480 nm generates the colored open form **4a** that reverts back to the cyclic form **3a** in the dark.<sup>13</sup> The formation of the cyclic isomers is the unprecedented phenomenon for 1,3-indandione derived merocyanines, which behave as typical donor – acceptor conjugated systems.<sup>14</sup> Moreover, we find that the photochemical properties of this compound and its analogs strikingly differ from the behavior of the hitherto known photochromic compounds. Here, we report on the synthesis, photochromic behavior, and X-ray structures of a series of derivatives of type **3**.

Derivatives **3** were prepared by condensation of bindone **1** with the commercially available *p*-substituted cinnamic aldehydes **2** in acetic anhydride or ethanol in 42 – 90% yields as reddish (**3a**, **b**) or yellow-orange (**3c**) crystals.



Scheme 1. Synthesis and photochromic conversions of derivatives **3**.

The X-ray crystal structures<sup>†</sup> were determined for derivatives **3a** – **3c** (Fig. 1) and confirmed their cyclic forms. There

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discoloration rate by another order of magnitude (Fig. 3b). In protic solvents coloration can be observed only under irradiation and the colored forms disappear within few seconds after the light source is off. Complete restoration of the **4c** spectrum takes about 6 h (in methylene chloride) and several days (in toluene). The colored forms rapidly disappear at elevated temperatures: in toluene, for instance, discoloration of **4a** is almost instantaneous at 90–100°C.

The MC forms are not photosensitive. Thus, **4a** cannot be generated by irradiation of **3a** at 642 nm in *n*-octane, since **3a** practically does not absorb at this wavelength. On the contrary, absorption of **4a** is still considerable here. Discoloration of **4a**, generated at 405 nm, under permanent irradiation at 642 nm, yielded the same discoloration rate constant as the dark measurements. The discoloration rate of **4c** in methylene chloride was not affected by light at 642 nm as well.

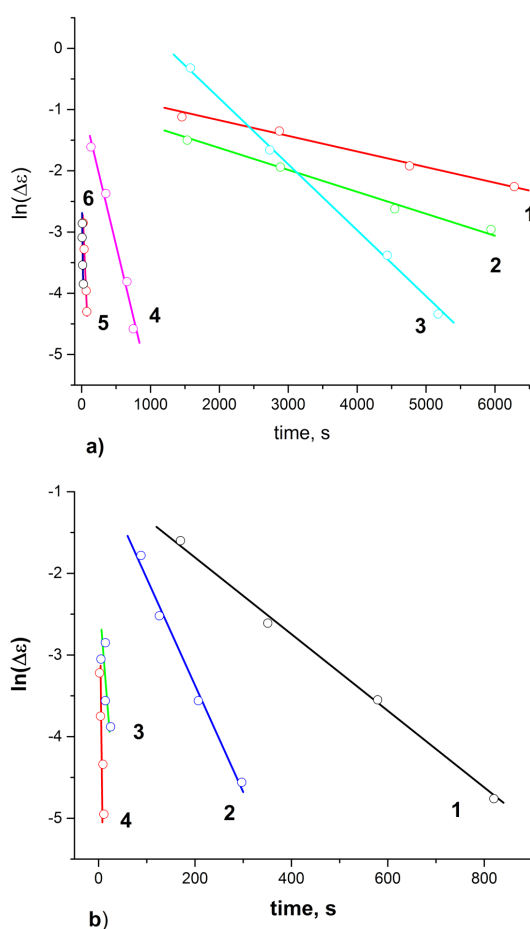


Fig. 3. Discoloration kinetics at 25°C of a) **4a** in different solvents ( $k \times 10^4 \text{ s}^{-1}$ ): 1 – cyclohexane (2.55), 2 – *n*-octane (3.33), 3 – toluene (10.8), 4 – THF (47), 5 – methylethyl ketone (258), 6 –  $\text{CH}_2\text{Cl}_2$  (744); b) **4a** and **4b** ( $k \times 10^4 \text{ s}^{-1}$ ): 1 – **4a** in THF (47), 2 – **4b** in THF (130), 3 – **4a** in  $\text{CH}_2\text{Cl}_2$  (744), 4 – **4b** in  $\text{CH}_2\text{Cl}_2$  (4810).

The photochemical inertness of the photo-induced isomers **4** is not unexpected, taking into account their behaviour as typical donor – acceptor conjugated systems with intramolecular charge transfer. Thus, strong positive solvatochromism of **4** (Fig. 4) confirms the presence of the

efficient conjugation observed also with their non-photochromic analogs<sup>14</sup>. The longest wavelength absorption band of derivative **4a** undergoes a 62 nm red shift in methylethyl ketone compared to the octane solution, confirming its nature as a charge transfer band. The electronic transitions of this kind are photochemically inert. For instance, the efficient conjugation between the electron donating and accepting moieties leads, in particular, to the loss of photochromism in fulgides.<sup>18</sup>

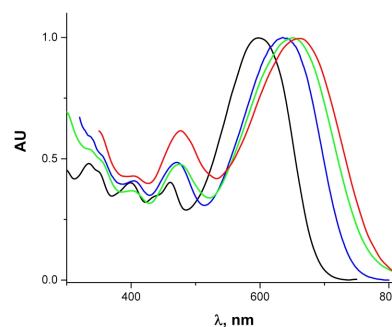


Fig. 4. Normalized absorption spectra of **4a** in *n*-octane (black), toluene (blue), THF (green) and methylethyl ketone (red).

Repeating irradiation at 532 nm and thermal discoloration of **3a – c** for 10 – 15 times did not reveal any noticeable fatigue features.

Quantum mechanical calculations† on the cyclic derivative **3a** and all four isomers of the MC form using B3LYP/6-31G(d,p)//B3LYP/6-31G(d,p) model chemistry showed that the most stable isomers are **4a** (*Z,E*), by 10.4 kcal/mol and **4'a** (*E,E*), by 7.0 kcal/mol (ZPE corrected) more stable than cyclic **3a**. *Z,Z*- and *E,Z*-isomers are the least stable, but still more stable than **3a** by 4.7 and 3.3 kcal/mol, respectively. Calculations in methylene chloride increased the energy difference by about 1 kcal/mol in favor of the MC isomers. Using the B3LYP/6-311G+(2d,p) basis set gave essentially the same results.

TD DFT calculations at the same model chemistry predicted that the first electronic transition within derivatives **3** (581 nm,  $f = 0.009$ ) is predominantly the HOMO – LUMO transition of the  $\pi, \pi^*$ -type (Fig. 5). Upon excitation, charge transfer occurs from the electron donating aryl groups toward the neighboring electron accepting fluorenyl and 1,3-indandionyl moieties. The low intensity longest wavelength absorption bands observed for derivatives **3** are thus *through-space charge transfer* bands in contrast to the MC-type isomers of **4** exhibiting intense *through-bond charge transfer* bands (549 nm,  $f = 0.7$ ). The degree of charge transfer can be roughly evaluated by the change in the NBO charges of the C=O groups. Thus, excitation leads to the increase in the sum of negative charges on the C=O groups by 0.19 e<sup>-</sup> (**3a**) and by 0.04 e<sup>-</sup> for (**4a**).

The mechanism of the photochemical reaction is currently under investigation in our lab. Our preliminary results strongly suggest that the photochromic behavior of derivatives **3** may

stem from at least partial retaining of the HBs, observed in solid, also in solution.

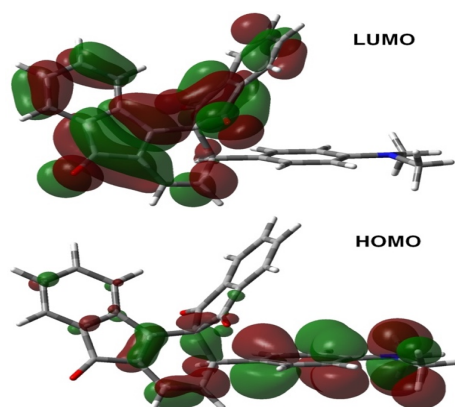


Fig. 5. HOMO and LUMO plots of **3a**.

Derivatives **3a** – **c**, as new photochromes, exhibit thus a number of unique features, such as sensitivity to both UV and visible light, the formation of the insensitive to light colored forms and a wide span of their thermal discoloration rates. The ability of these compounds to react with green and red light and strong dependence of the discoloration rates on the polarity of surroundings suggest their possible use for biomedical studies.

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## Notes and references

† Crystallographic data for the structures reported in this paper have been deposited with the Cambridge Crystallographic Data Centre: **3a** CCDC 1436240; **3b** CCDC 1436241 and **3c** CCDC 1436242.

- A. V. Kulinich and A. A. Ishchenko, *Russ. Chem. Rev.*, 2009, **78**, 141.
- S. R. Marder, B. Kippelen, A. K.-Y. Jen and N. Peyghambarian, *Nature*, 1997, **388**, 845; A. Teshome, A. J. Kay, A. D. Woolhouse, K. Clays, I. Asselberghs and G. J. Smith, *Optical Mater.*, 2009, **31**, 575; U. Lawrentz, W. Grahn, K. Lukaszuk, C. Klein, R. Wortmann, A. Feldner and D. Scherer, *Chem. Eur. J.*, 2002, **8**, 1573.
- A. Mishra, M. K. Fischer and P. Bäuerle, *Angew. Chem. Int. Ed.*, 2009, **48**, 2474; A. Zitzler-Kunkel, M. R. Lenze, N. M. Kronenberg, A.-M. Krause, M. Stolte, K. Meerholz and F. Würthner, *Chem. Mater.*, 2014, **26**, 4856; T. Meyer, D. Ogermann, A. Pankrath, K. Kleinermanns and T. J. Müller, *J. Org. Chem.*, 2012, **77**, 3704.
- L. Huang, M. Stolte, H. Bückstümmer and F. Würthner, *Adv. Mater.*, 2012, **24**, 5750.
- A. Toutchkine, W.-G. Han, M. Ullmann, T. Liu, D. Bashford, L. Noodleman and K. M. Hahn, *J. Phys. Chem. A*, 2007, **111**, 10849; M. Li, X. Wu, Y. Wang, Y. Li, W. Zhu and T. D. James, *Chem. Commun.*, 2014, **50**, 1751; C. J. MacNevin, D. Gremyachinskiy, C.-W. Hsu, L. Li, M. Rougie, T. T. Davis and K. M. Hahn, *Bioconjugate chemistry*, 2013, **24**, 215.
- Organic Photochromic and Thermochromic Compounds. Crano, J. C.; Guglielmetti R. J. Eds. Plenum Press, N.-Y. 1999; V. I. Minkin, *Chem. Rev.*, 2004, **104**, 2751; V. I. Minkin, *Russ. Chem. Rev.*, 2013, **82**, 1.
- V. A. Lokshin, A. Samat and A. V. Metelitsa, *Russ. Chem. Rev.*, 2002, **71**, 893; J. D. Hepworth and B. M. Heron, in *Functional Dyes*, ed. S.-H. Kim, Elsevier, Amsterdam, 2006, p. 85.
- O. Brede, L. Goebel, L. and T. Zimmermann, *J. Phys. Chem. A*, 1997, **101**, 4103; L. Sanguinet, J. L. Pozzo, V. Rodriguez, F. Adamietz, F. Castet, L. Ducasse and B. Champagne, *J. Phys. Chem. B*, 2005, **109**, 11139.
- Y. Shiraishi, M. Itoh and T. Hirai, *Phys. Chem. Chem. Phys.*, 2010, **12**, 13737; J. Whelan, D. Abdallah, K. Piskorz, J. T. C. Wojtyk, J. M. Dust, J.-M. Nunzi, S. Hoz and E. Buncel, *Phys. Chem. Chem. Phys.*, 2012, **14**, 13684; Z. Tian, R. A. Stairs, M. Wyer, N. Mosey, J. M. Dust, T. M. Kraft and E. Buncel, *J. Phys. Chem. A*, 2010, **114**, 11900.
- D. G. Patel, M. M. Paquette, R. A. Kopelman, W. Kaminsky, M. J. Ferguson and N. L. Frank, *J. Am. Chem. Soc.*, 2010, **132**, 12568; Y. Shiraishi, M. Itoh and T. Hirai, *Tetrahedron Lett.*, 2011, **52**, 1515; S. Aiken, K. Booth, C. D. Gabbutt, B. Mark Heron, C. R. Rice, A. Charaf-Eddin and D. Jacquemin, *Chem. Commun.*, 2014, **50**, 7900.
- G. Copley, J. G. Gillmore, J. Crisman, G. Kodis, C. L. Gray, B. R. Cherry, B. D. Sherman, P. A. Liddell, M. M. Paquette, L. Kelbaskas, N. L. Frank, A. L. Moore, T. A. Moore and D. Gust, *J. Am. Chem. Soc.*, 2014, **136**, 11994; P. Peng, P. Strohhecker and Y. Liao, *Chem. Commun.*, 2011, **47**, 8575; J. Buback, P. Nuernberger, M. Kullmann, F. Langhojer, R. Schmidt, F. Würthner and T. Brixner, *J. Phys. Chem. A*, 2011, **115**, 3924; S. Helmy, F. A. Leibfarth, S. Oh, J. E. Poelma, C. J. Hawker and J. Read de Alaniz, *J. Am. Chem. Soc.*, 2014, **136**, 8169.
- A. A. Beharry, O. Sadovski and G. A. Woolley, *J. Am. Chem. Soc.*, 2011, **133**, 19684; D. Bléger, J. Schwarz, A. M. Brouwer and S. Hecht, *J. Am. Chem. Soc.*, 2012, **134**, 20597; S. Samanta, A. A. Beharry, O. Sadovski, T. M. McCormick, A. Babalhavajji, V. Tropepe and G. A. Woolley, *J. Am. Chem. Soc.*, 2013, **135**, 9777; Y. Yang, R. P. Hughes and I. Aprahamian, *J. Am. Chem. Soc.*, 2014, **136**, 13190.
- B. Zinger, P. Schaer, G. Berkovic, G. Meshulam, Z. Kotler, L. Shapiro, R. Mazor and V. Khodorkovsky, *SPIE Proceedings*, 1997, **3135**, 71.
- G. Meshulam, G. Berkovic, Z. Kotler, A. Ben-Asuly, R. Mazor, L. Shapiro and V. Khodorkovsky, *Synthetic Metals*, 2000, **115**, 219; A. Arjona-Esteban, J. Krumrain, A. Liess, M. Stolte, L. Huang, D. Schmidt, V. Stepanenko, M. Gsänger, D. Hertel, K. Meerholz and F. Würthner, *J. Am. Chem. Soc.*, 2015, **137**, 13524; J. Francos, S. E. García-Garrido, J. Borge, F. J. Suárez and V. Cadierno, *RSC Adv.*, 2016, **6**, 6858.
- Desiraju, G. R.; Steiner, T. *The Weak Hydrogen Bond in Structural Chemistry and Biology*. Oxford University Press, New York, 1999; S. Horowitz and R. C. Triebel, *J. Biol. Chem.*, 2012, **287**, 41576.
- F. Orozco, B. Insuasty, J. N. Low, J. Cobo and C. Glidewell, *Acta Cryst. C*, 2008, **64**, o162.
- M. Sigalov, A. Vashchenko and V. Khodorkovsky, *J. Org. Chem.*, 2005, **70**, 92; M. Sigalov, P. Krief, L. Shapiro and V. Khodorkovsky, *Eur. J. Org. Chem.*, 2008, 673; C. Niebel, V. Lokshin, M. Sigalov, P. Krief and V. Khodorkovsky, *Eur. J. Org. Chem.*, 2008, 3689.
- P. J. Darcy, H. G. Heller, S. Patharakorn, R. D. Piggott and J. Whittall, *J. Chem. Soc. Perkin Trans. I*, 1986, 315.