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Correlation between photoluminescence and photoelectrochemical properties of SrHPO₄/BaHPO₄/FTO anode material

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Abstract

A new anode consisting of BaHPO₄/SrHPO₄ (SrP/BaP) micro-structured materials has been fabricated through two-step electrodeposition technique. Structural, vibrational, morphological, and optical properties of the anode films were investigated employing X-ray diffraction, Raman spectroscopy, scanning electron microscopy (SEM) and UV–Vis absorption spectroscopy, respectively. Photoluminescence (PL), and photo-electrocatalytic (PEC) activities of the samples were investigated and compared with those of the pure BaHPO₄ (BaP) and SrHPO₄ (SrP) samples. **The PEC performance was of the order of 71 % within 25 min, which was less than both BaP (99%) and SrP (92%) in the same experimental conditions. These results** reveal that BaP thin films combined with SrP microstructures remarkably changed the PL signals, and in turn significantly enhanced the PL activities, while the PEC has been lowered down. This was attributed to the photo-induced electron–hole pairs recombination as electrochemical impedance spectrum and photo-current response have shown. Based on these results, it is believed that the SrP/BaP film behaves as a sensitizer in the system and not as a photo-electrocatalyst.

Keywords

SrHPO₄/BaHPO₄; photo-electrocatalysis; photoluminescence; charge transfer; thin films.

1

2 **1. Introduction**

3 The luminescence spectrophotometry is a powerful tool for understanding the
4 photo/electrocatalysis by studying the electronic structure, optical and photochemical
5 properties of semiconductor materials. It is also used to obtain information about surface
6 defects, oxygen vacancies, photo-induced charge carrier separation and recombination
7 processes ... etc. [1,2]. To understand the luminescence origin, it is important to keep in mind
8 that the creation of electron-hole pairs within semiconductor materials generally leads to two
9 extreme mechanisms. One is the recombination of these charge carriers while in the second
10 case it is their efficient separation in a redox process, which occurs. The first mechanism is
11 usually accompanied with a release of a certain amount of chemical energy which will further
12 transform to heat or to light energy [1]. These latter result in a luminescence emission of the
13 semiconductor material called PL phenomenon of the semiconductor. The second mechanism
14 is usually beneficial for photocatalytic applications as the recombination of charge carriers
15 will be retarded or even avoided, which improves the efficiency of the process [3,4].
16 However, in some cases, it was found that both photo/electrocatalytic activity and
17 photoluminescence (PL) of semiconductor photocatalysts can increase likewise [5,6].

18 Recently, thanks to the increase in the development of science and technology, the inorganic
19 semi-conductor materials with different morphologies and physico-chemical properties such
20 as high surface areas, optical, electrical, and surface properties have shown a promising
21 perspective in many fields, such as catalysis, coatings, and sensors [7–9]. Phosphate-based
22 catalysts are one of the important classes of semiconductors which have been paid intense
23 attention because of their excellent properties, such as the large band gap, high absorption of
24 PO_4^{3-} in UV-Vis region, moderate phonon energy, high thermal as well as their high chemical
25 stability. Therefore, they can be used in various applications such as supercapacitors, sensors,
26 luminescence, Li-ion batteries, separation of radioactive ions, and photo/electrocatalysis [10–
27 19] .

28 Among phosphate-based catalysts, strontium hydrogen phosphate (SrHPO_4) and barium
29 hydrogen phosphate (BaHPO_4) have been investigated for a long time by many researchers as
30 photocatalysts, surface conditioner, in batteries, fuel cells, proton conductor, for flame
31 proofing, adsorption of heavy metals and dyes and thermal cathodes [20–25]. **Thus, in view of**

1 the above-mentioned studies on BaP and SrP based materials and their wide applications, it is
2 thought worthwhile to study their composite material.

3 Because neither the PEC and PL properties of the SrHPO₄/BaHPO₄ composite nor the PL
4 performance of SrHPO₄ and BaHPO₄ alone have not been studied, the correlation between the
5 two properties as well, this paper is considered as a first attempt to study these properties.

6 For this reason, in the present work, we fabricated SrP/BaP by two-step electrochemical
7 deposition for the first time to our knowledge. After a detailed morphological and structural
8 characterization, the correlation between the photoluminescence (PL) and the
9 photoelectrochemical properties was studied by measuring PL and PEC. The goal being to
10 understand the suppressed radiative recombination and the charge separation of
11 photogenerated carriers in the SrP/BaP films.

12 **2. Materials and methods**

13 **2.1. Materials**

14 Strontium (III) nitrate Sr(NO₃)₂, barium nitrate Ba(NO₃)₂, ammonium dihydrogen phosphate
15 (NH₄)H₂PO₄, were provided from Fluka while Rhodamine B (RhB) was from Sigma–Aldrich.
16 All these materials were used as received without further purification. F-doped SnO₂ (FTO)
17 glass was purchased from SOLEMS S.A. France. Its sheet resistance is 6-8 Ω/square.

18 **2.2. Preparation of BaP–SrP anode, PEC and PL procedures**

19 The BaP–SrP anode was prepared by a two-step electrodeposition method on FTO substrates.
20 Prior to the deposition, FTO substrates were ultrasonically cleaned in acetone, and deionized
21 water, for 15 min sequentially. Firstly, an aqueous electrocatalytic solution composed by 0.03
22 M of strontium nitrate solution (50 mL) and 0.02 M of ammonium dihydrogen phosphate
23 solution (50 mL) was prepared. The FTO substrates served as the working electrode, graphite
24 rod was used as the counter electrode and the saturated calomel electrode (SCE) as the
25 reference electrode. The electrodeposition procedure was performed at a temperature of 23°C
26 and under a constant current density of -1 mA cm^{-2} relative to the counter electrode.
27 Consequently, the electrochemical deposition of BaP on the SrP films was performed in an
28 aqueous solution composed of 0.03 M of barium nitrate solution (50 mL) and 0.02 M of
29 ammonium dihydrogen phosphate solution (50 mL). In this case the FTO/SrP films served as
30 the working electrode. The applied current density was controlled at -0.5 mA cm^{-2} vs the
31 counter electrode. The deposition procedure is illustrated in figure 1 below.

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Figure 1: Schematic diagram of the deposition procedure of the SrP/BaP films.

The PEC degradation was carried out using the Rhodamine B as a model dye [20,26,27]. The process was achieved in an electrochemical cell consisting of three electrodes, SrP-BaP film as the photoanode, ECS as a reference electrode while the counter electrode was a carbon graphite. The reaction solution (100 mL) contains RhB solution and NaCl as supporting electrolyte. Samples are collected at specific time intervals for the determination of residual RhB concentration. The UV light source was a Xenon lamp (250 W).

The color removal efficiency was calculated according to the following equation:

$$\% \text{ color removal} = 100 \times ((C_0 - C_t) / C_0) \quad (1)$$

Where C_0 and C_t are the concentration of RhB solution before the electrolysis and at each time “t” during the electrolysis, respectively.

The photoluminescence experiments were performed by using a Horiba JobinYvon LabRam HR800 spectrometer. This later is equipped with an argon-ionized laser as a source of excitation with a wavelength of 633 nm (1.95 eV) and whose power was fixed at 5 μ W. The control and acquisition of spectra was achieved using LabSpec software. The measurements were recorded in the spectral range between 400 and 1200 nm with an acquisition time of the spectra of 5 seconds.

2.3. Characterization

X-ray diffraction (XRD) using Bruker D8 Advance Twin diffractometer ($\lambda_{\text{Cu-K}\alpha}$ = 0.154 nm) and Raman spectroscopy (Horiba Jobin-Yvon HR800 LabRam system) were used for SrP-BaP thin films structural analysis while Jeol JSM-6460LAV SEM operating at 20kV and equipped with an energy-dispersive X-ray spectrometer (EDS) was used to determine the morphology and the chemical composition of these films. UV-Visible spectrophotometry was employed to analyze the optical and pho-electrocatalytic behavior of SrP-BaP thin films.

2.4. Experimental design

Optimization experiments were carried out using response surface methodology (RSM), and in particular Central Composite Design (CCD) in which four factors were optimized [28]. The

1 experimental factors studied are initial NaCl concentration (X_1), applied current density (X_2),
2 initial RhB concentration (X_3) and reaction time (X_4). These factors were varied over five
3 levels between $-\alpha$ and $+\alpha$ at the determined ranges (Table SI 1). Based on the preliminary
4 experiments and the experimental set-up, the choice of factors and their values were
5 determined. The experimental data were fitted to second-order polynomial model according to
6 the following equations:

$$N=2^k+2k+N_0 \quad (2)$$

$$Y = b_0 + \sum_{i=1}^n b_i X_i + \sum_{i=1}^n b_{ii} X_i^2 + \sum_{i=1}^{n-1} \sum_{j=i+1}^n b_{ij} X_i X_j \quad (3)$$

9 Where N is the total number of experiments, k is the number of parameters and N_0 is the
10 number of central points. Y is the degradation efficiency of RhB; X_i and X_j are coded values
11 of the factors (i and j range from 1 to k); b_0 is the intercept coefficient of model; b_j , b_{ii} , and b_{ij}
12 are interaction coefficients of linear, quadratic, and the second-order terms, respectively; n is
13 the number of independent parameters ($n=4$ in this study).

14 The analysis of variance (ANOVA) was performed in order to determine the individual linear,
15 quadratic and interaction regression coefficient, this was performed using NONEMROD 2007
16 software. The quality of the fit of polynomial model was expressed by the value of correlation
17 coefficient (R^2) and adjusted coefficient of determination (R^2_{adj}) as well as the ANOVA
18 results [28].

19

20 **3. Results and discussion**

21 **3.1. X-ray diffraction and Raman analysis**

22 The XRD patterns were recorded to confirm the structure, phase and crystallinity of SrP-BaP
23 as shown in Fig. 2.a The β -SrHPO₄ hexagonal phase is observed which is in well agreement
24 with the JCPDS file no. 12-0368 (labeled as “●”). The XRD peaks for BaHPO₄ match the
25 standard peak value with the orthorhombic phase of BaHPO₄ (JCPDS file no. 01-072-1370)
26 (labeled as “*”). It is worth to mention that BaP patterns are relatively more intense than those
27 of SrP which is obvious because the BaP was deposited on SrP. BaP also filled the pores that
28 may be present during the formation of SrP. The detected additional peaks correspond to the
29 standard FTO peaks. It is noteworthy that no other diffraction peaks were observed. This
30 confirms the synthesis of the SrP/BaP without any impurities.

31 The Raman spectrum is shown in Figure 2.b which is composed of the
32 following wavenumbers. The bands at 1055 and 982 cm^{-1} are assigned to antisymmetric

1 stretching internal modes of the phosphate internal modes (ν_3 ($A1g + Eg$)), while the band at
2 943 cm^{-1} for symmetric stretching of P-O (ν_1 ($A1g$)). The two bands located at 580 and 533
3 cm^{-1} are due to the deformation bending of O-P-O (ν_4 ($A1g + Eg$)), while the intense band at
4 401 cm^{-1} is attributed to the bending vibration (ν_2 (Eg)). The external modes deriving from
5 the vibrations of PO_4 , Sr-O and Ba-O at 207, 145 and 105 cm^{-1} [29,30]. The Raman results
6 further confirmed the formation of SrP/BaP anode material.

7

8 **Figure 2:** a) The X-ray diffraction patterns of FTO substrate and SrP/BaP b) Raman spectrum
9 of SrP/BaP anode material.

10

11 **3.2. SEM and EDS analysis**

12 Figure 3 presents the SEM images of SrP (a), BaP (c and SrP-BaP (b), it is clear that all of the
13 films possess homogeneous surface morphology and without cracks. SrP shows flower-like
14 morphology with the average diameter of flowers in the range of $10\text{ }\mu\text{m}$ while BaP shows a
15 parallelepiped morphology with the diameter and length in the range of $4\text{ }\mu\text{m}$ and $16\text{ }\mu\text{m}$,
16 respectively. Additionally, the surface of SrP-BaP films was rougher than both the SrP and
17 BaP films surface. We inferred that the formation of granular structure on the surface of film
18 should be responsible for the increase of roughness since there is a slight agglomeration by
19 mixtures of both SrP and BaP particles when BaP particles were introduced.

20 Energy-dispersive spectrometry (EDS) analysis of SrP-BaP was performed to further confirm
21 the elemental analysis. As shown in table 1, the exclusive occurrence of the Sr, Ba, O, and P
22 elements suggested the high purity of the films.

23 It is worth to mention that the areas delimited in the figure 4.c do not correspond to the given
24 images of SrP (in figure 4.a) and BaP (in figure 4.c) but for the objective to show the
25 existence of the two observed morphologies.

26

27 **Figure 3:** SEM images of (a) SrP, (b) SrP/BaP and (c) BaP thin films

28 **Table 1:** EDS analysis of SrP-BaP thin films

29

30 **3.3. Optical properties and conductivity type**

1 UV–vis diffuse reflectance spectroscopy for the synthesized samples was performed to
2 monitor the optical response of the synthesized SrP-BaP thin film. The results are shown in
3 Fig. 4a. The optical band gap energy was calculated by Tauc plots of $(\alpha h\nu)^2$ versus $h\nu$ [31,32],
4 the estimated band gap of the obtained films was found to be 3.85 eV. The results were in a
5 good agreement with literature data [25,33].

6 The Mott-Schottky analysis was performed to identify the type of conductivity for SrP/BaP
7 (Fig.4b). Indeed, linear region of $1/C^2$ (C is the capacitance) relative to the applied potential
8 was fitted with the Mott-Schottky equation. The positive slope of the line suggests n-type
9 conductivity of SrP/BaP anode [34]. The corresponding intercept to the x -axis provides the
10 flat-band potential, which can be taken as a reference for Fermi level of the semiconductor.

11

12 **Figure 4:** a) Estimation of band gap energy, b) conductivity type of the SrP/BaP anode film.

13

14 **3.4. Experimental design and data analysis**

15 Table SI-2 presents the conditions of 30 experimental runs, and the experimental response in
16 terms of RhB degradation efficiency. According to the data in Table SI-2 and SI-3, the
17 experimental design for % degradation of RhB in terms of coded factors was modelled by a
18 second-order polynomial equation given below (Eq. 4):

19 % , Degradation rate of RhB = $229.27 - 23.72 X_1 - 14.23 X_2 - 30.83 X_3 + 45.85 X_4 - 42.89 X_1^2 -$
20 $37.82 X_2^2 - 30.21 X_3^2 - 30.21 X_4^2 + 0.97 X_1 X_2 + 7.75 X_1 X_3 + 7.75 X_1 X_4 - 0.97 X_2 X_3 + 0.97$
21 $X_2 X_4 - 17.43 X_3 X_4$ (4).

22 Results from the CCD experiments and the **surface response** counter plots in figure SI-1 and
23 Table SI-3, the degradation of RhB depends on all the four studied variables, the linear terms
24 and the quadratic terms having either negative or positive effects. However, the interaction of
25 NaCl concentration and applied current density, the initial dye concentration and the applied
26 current density are not significant as p -value for all these two terms is greater than 0.05. All
27 the other terms are significant as p -values for them are less than 0.05% [35]. However, the
28 non-significant terms were not eliminated from the model because it would reduce the
29 dimensionality and thus the accuracy of the response surface. Based on the values of
30 independent variables coefficients in the regression model, the order in which the studied

1 variables affect the response as follows: reaction time (X_4) > RhB initial concentration
2 (X_3) > initial NaCl concentration (X_1) > applied current density (X_2).

3 Finally, to test the significance of the model, the ANOVA was carried out and the obtained
4 results are given in Table SI-3. A very low probability value ($P_{\text{model}} < 0.0001$) indicated that
5 the model obtained was highly significant [36]. Furthermore, the R^2 test has been also used
6 for the validation of the mathematical model. The R^2 of the quadratic model was 0.94.
7 Adjusted R^2 (Adj- R^2), which is also a measure of goodness of a fit, was 0.89, indicating a
8 satisfactory adjustment of the polynomial model to the experimental data. The final validation
9 of the experimental model of CCD is confirmed if the selected model provides adequate
10 approximation of the real system. The correlation between experimental and predicted values
11 in the optimal conditions is given in Table 2. Based on the obtained results, it can be
12 concluded that there is a good agreement between the experimental and predicted values. The
13 degradation rate of RhB was 54 %, which is in close approximation to the predicted value
14 52.8 %. The results show that the proposed model used in this study is a suitable and usable to
15 describe the aqueous electrodegradation of RhB onto SrP/BaP electrodes.

16 **Table 2:** Optimum values of the factors studied for optimal RhB degradation.
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19 **3.5. Photo-electrocatalytic properties**

20 To investigate the PEC activity of SrP/BaP electrode, experiments using the optimal
21 conditions found by the RSM (shown in Table 2) were performed and the obtained results are
22 given in Fig. 5. Initially, the thin films were immersed in RhB dye solution in the dark for 60
23 minutes and the concentration of the dye was measured in order to achieve the adsorption of
24 dye onto the electrode, the obtained results revealed that a very negligible decrease in the
25 concentration of the dye was observed (1 %). The photolysis of RhB dye was also
26 investigated and the concentration did not decrease, only 8% of degradation was obtained
27 indicating the photostability of RhB dye during the reaction time of experiments. The
28 photocatalysis experiment was conducted without applying a current density and after 25 min,
29 14 % of initial dye concentration was eliminated. The electrocatalysis experiment were
30 performed by applying $13.6 \text{ mA}\cdot\text{cm}^{-2}$ without light illumination; 54 % degradation efficiency
31 was obtained. On the other hand, the UV light illumination with applying a current density
32 have led to a superior degradation performance of the order of 71 % within 25 min. This

1 superior PEC activity can be attributed to the synergistic effect between photocatalysis and
2 electrocatalysis.

3 To explain the synergic effect shown, when the SrP/BaP is subjected to an applied current
4 density in combination with UV light, the separation of electrons-holes pairs is favored as the
5 photogenerated electrons can travel easily from the working to the counter electrode via the
6 external circuit. This also increases the amount of charge carriers being available in the
7 system which therefore increases the number of active radicals and thus promotes the
8 photoelectrodegradation efficiency.

9

10 However, comparing BaP and SrP PEC efficiencies in the same conditions, the SrP/BaP
11 shows less performance as shown in figure 6. High recombination rates are believed to be
12 behind the observed phenomena as it will be explained by photoluminescence and impedance
13 studies, and thus limits the efficiency process.

14

15 **Figure 5:** Comparison of adsorption, photocatalysis, electrocatalysis and photo-
16 electrocatalysis performances of SrP-BaP electrode. Experimental conditions: reaction time:
17 25 min, [NaCl]=0.18 M, current density = 13.6 mA.cm⁻², and [RhB]=6 M.

18

19 **Figure 6:** Comparison of PEC process of different electrodes under the same conditions.

20

21 3.6. PL spectra analysis

22 SrHPO₄ and BaHPO₄ are one kind of commonly used luminescent host materials. For
23 example, SrHPO₄ doped with Eu³⁺, Ce³⁺, Tb³⁺ have shown important luminescent properties
24 under the excitation of UV light [37]. However, these two materials have been doped with
25 luminescent centers such as rare earth elements which are typically expensive. Thus, much
26 effort should be devoted to explore self-activated luminescent materials that do not contain
27 expensive elements. Therefore, the photoluminescence properties of SrP, BaP and SrP/BaP
28 films have been investigated and shown in figure 7 below. The PL curves of pure SrP and
29 pure BaP are measured with the same PL system and identical settings for the purpose of
30 comparisons. BaP exhibits a broadband green emission from 350 to 800 nm and peaking
31 around the wavelength of 522 nm. While SrP shows important luminescent properties than

1 BaHPO₄. Indeed, a broad emission band extending from 380 to 800 nm and peaking around
2 522 and 646 nm corresponding to green and red emissions, respectively, is observed. This
3 broad band suggests that the emission process is typical of a multiphonon or multilevel
4 process, i.e., a system in which the relaxation occurs by means of several paths and involving
5 the participation of numerous states [38]. Compared to the pure BaP and SrP, the SrP/BaP
6 film shows a great enhancement in its light emission indicating that the luminescence
7 properties for the SrP/BaP are enhanced by BaP coating on SrP. This can be explained by the
8 existence of larger number of defects of metal atom and oxygen vacancies, compared to BaP
9 and SrP separately, during its crystal growth process [39].

10 Moreover, the higher recombination rate of photo-generated electrons and holes are also
11 responsible for the luminescence. Indeed, the obtained results from PEC studies indicated that
12 there was not charge carriers recombination suppression. Such results were reported
13 elsewhere [40,41].

14

15 **Figure 7:** PL properties of the various films elaborated

16

17 **3.7. EIS analysis and photo-current response**

18 Using electrochemical impedance spectroscopy (EIS), Nyquist plots of BaP, SrP and
19 SrP/BaP were carried out to give further evidence on the recombination rate. These plots are
20 displayed in the figure 8a. All the three catalysts exhibit semicircles in the studied frequency
21 region. The arc radius on the Nyquist plot of BaP is smaller than that of SrP, while that of
22 SrP-BaP is bigger than those of SrP and BaP indicating that the charge transfer resistance is
23 higher in the BaP/SrP film. This also indicated that a higher charge recombination and a less
24 efficient charge separation were observed for the SrP/BaP films.

25 To shed more light on the separation efficiency of electrons and holes. Photocurrent
26 measurement was used to investigate the excitation and transfer of photo-generated charge
27 carriers in SrP/BaP anode material. In fact, the photocurrent is formed mainly by transferring
28 photogenerated electrons, formed while exciting the photocatalyst by UV light, to the counter
29 electrode; the higher photocurrent, usually of the order of Ampere, indicates more effective
30 separation and longer lifetime of the photogenerated electrons. Figure 8b shows the
31 photocurrent density versus the irradiation time curve of SrP/BaP samples under several on-

1 off illumination cycles. The photocurrent density obtained was almost $0.06 \mu\text{A}\cdot\text{cm}^{-2}$ indicating
2 a higher recombination and a less efficient separation of photo-generated electron-hole pairs
3 occurred at the interface between SrP and BaP in the SrP/BaP anode material.

4
5 **Figure 8:** a) Electrochemical impedance spectroscopy of SrP, BaP and SrP/BaP, b) Photo-
6 current response of the SrP/BaP anode material.

7 **Conclusion**

8 In summary, SrP/BaP anode was synthesized using two-step electrodeposition. The study of
9 optical properties has shown a band gap of 3.9 eV and n-type conductivity. The
10 morphological study shows the presence of the BaP particles over the SrP flowers while XRD
11 and EDS analysis have proved the existence of two phases without any impurities. The PEC
12 activity towards RhB degradation has been optimized using RSM in combination with the
13 CCD matrix. Although the synergic effect being generated by combining UV light with
14 current density, SrP/BaP exhibits relatively high luminescence activity and less PEC
15 performance in degrading RhB compared to pure BaP and SrP. The obtained results were
16 explained by charge transfer process for the SrP/BaP thin films. It was found that the
17 photogenerated charge carriers show high recombination rates as the photocurrent response
18 had revealed. The EIS as well has shown that the composite is characterized by a high charge
19 transfer resistance. Overall, the findings of this study revealed that SrP/BaP anode film can
20 be used to fabricate optical devices especially high-efficiency green and red light-emitting
21 devices.

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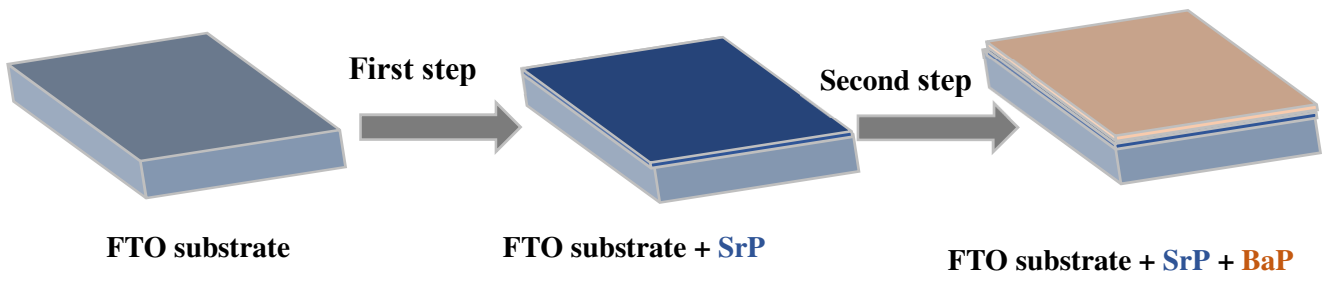


Figure 1: Schematic diagram of the deposition procedure of the SrP/BaP films.

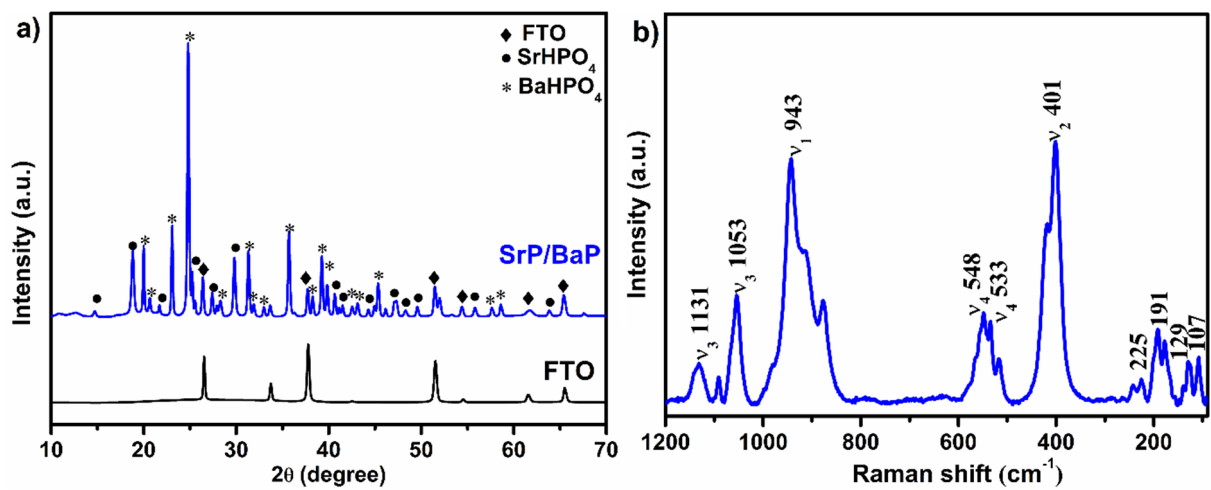


Figure 2: a) The X-ray diffraction patterns of FTO substrate and SrP/BaP b) Raman spectrum of SrP/BaP anode material.

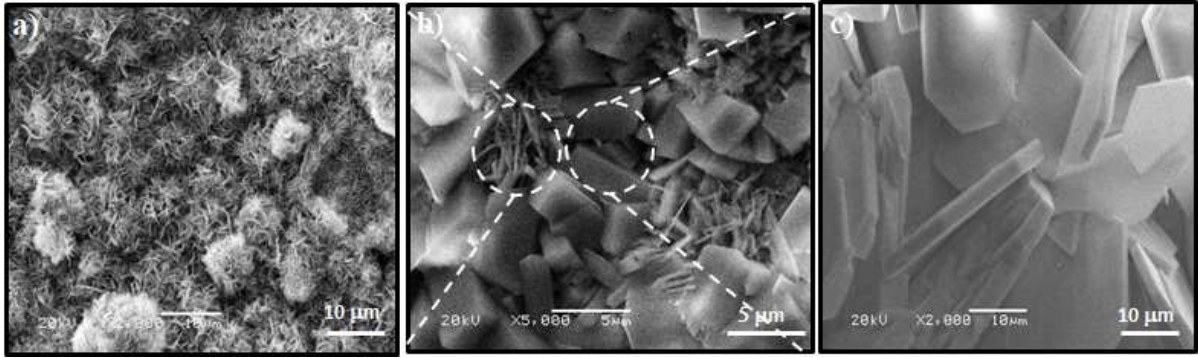


Figure 3: SEM images of (a) SrP, (b) SrP/BaP and (c) BaP thin films

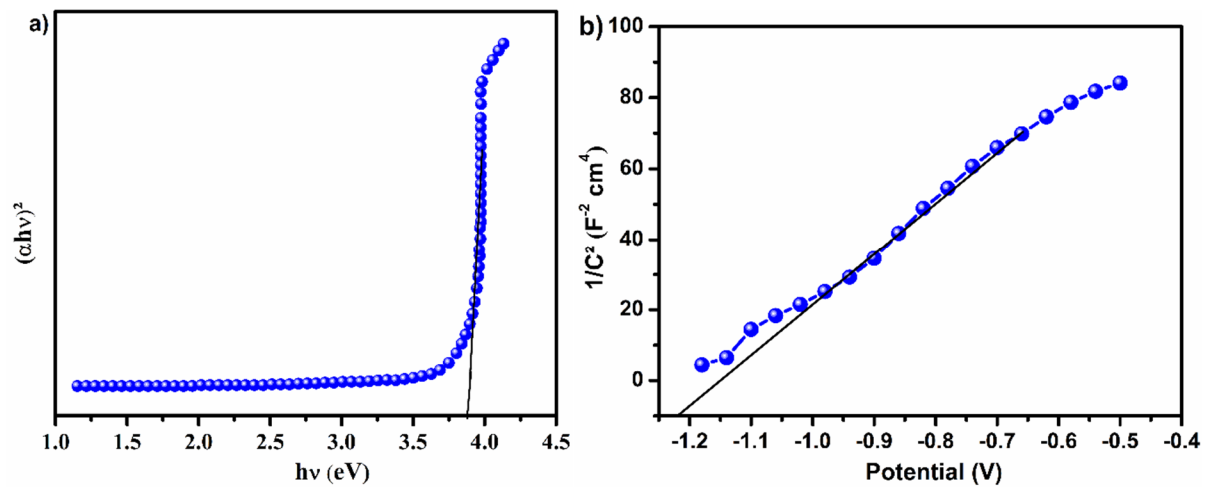


Figure 4: a) Estimation of band gap energy, b) conductivity type of the SrP/BaP anode film.

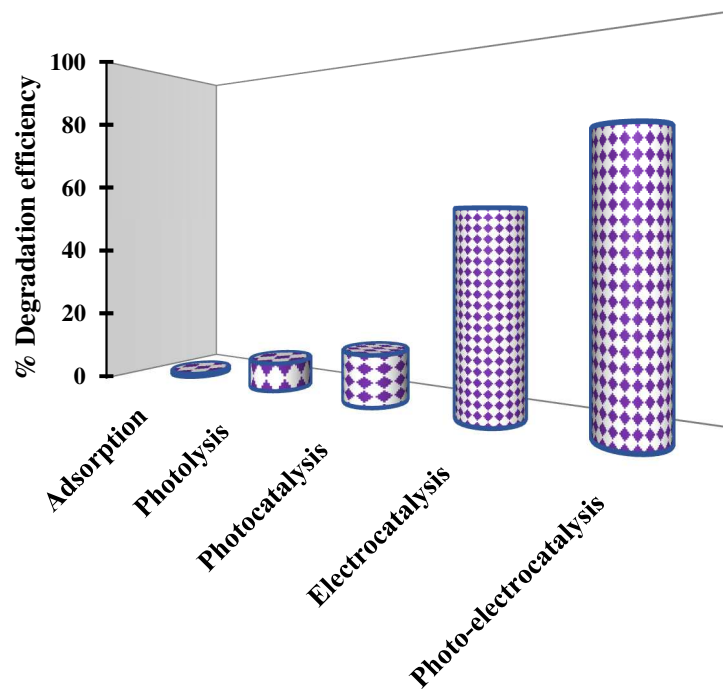


Figure 5: Comparison of adsorption, photocatalysis, electrocatalysis and photo-electrocatalysis performances of SrP-BaP electrode. Experimental conditions: reaction time: 25 min, $[\text{NaCl}] = 0.18 \text{ M}$, current density = $13.6 \text{ mA}\cdot\text{cm}^{-2}$, and $[\text{RhB}] = 6 \text{ M}$.

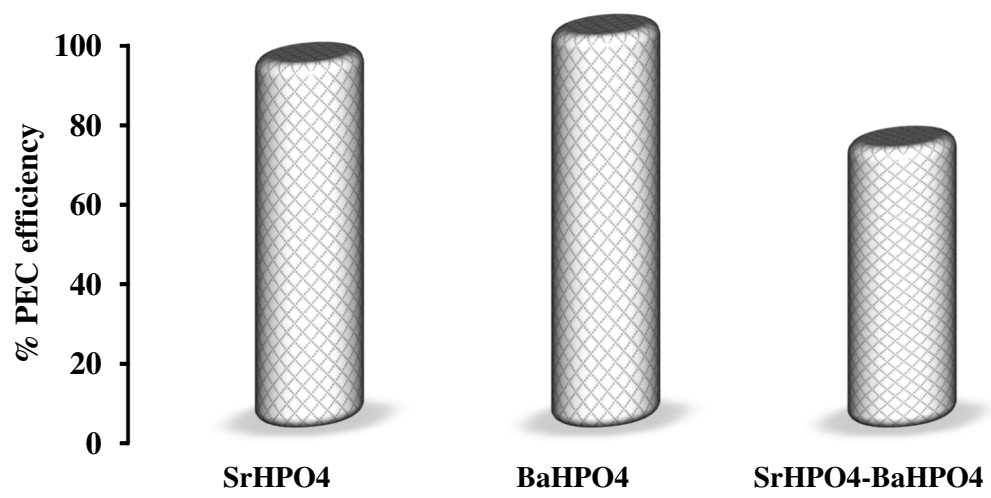


Figure 6: Comparison of PEC process of different electrodes under the same conditions.

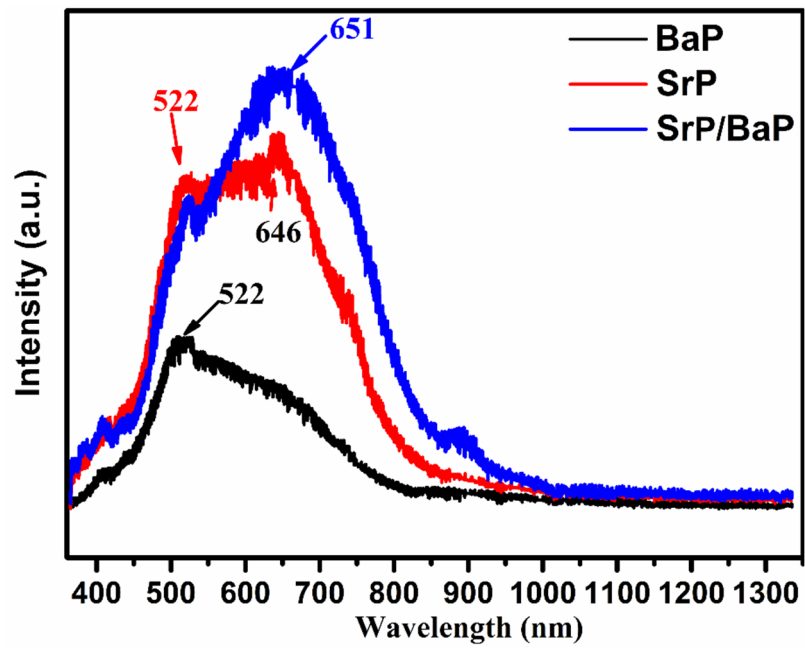


Figure 7: PL properties of the various films elaborated

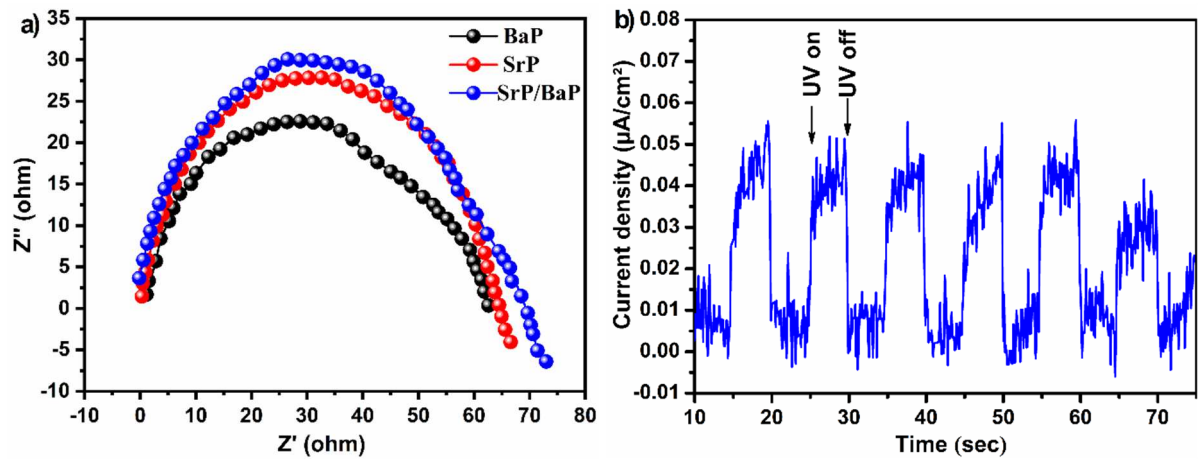


Figure 8: a) Electrochemical impedance spectroscopy of SrP, BaP and SrP/BaP, b) Photocurrent response of the SrP/BaP anode material.

Table 1: EDS analysis of SrP-BaP thin films

Element	At%	mass%
O K	47.9	16.1
P K	23.5	15.4
Sr L	13.7	25.3
Ba L	14.9	43.2
Total	100.00	100.00

Table 2: Optimum values of the factors studied for optimal RhB degradation.

Factor	Optimum values	Experimental Field	
		% found by the model	% found by experience
[NaCl] (mol/l): X₁	0.18		
Current density (mA/cm ²): X₂	13.60	52.80 %	54 %
[RhB ₀] (mol/l): X₃	6		
Time (min): X₄	25.00		

