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HEAT TRANSFER IN SALT CAVERNS

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Abstract

Several aspects of the thermodynamic behavior of salt caverns containing brine, oil, natural gas, air or hydrogen are discussed. It is shown that natural convection always appears in a liquid-filled cavern. This is not true in the case of gas caverns. Gas type, brine temperature at the cavern bottom and water vapor content are influential. Equations are obtained for the thermodynamic behavior of the fluids stored in a cavern during injection-withdrawal. Thermal evolution of a liquid-filled cavern at rest is slow; rapid injection and withdrawal are adiabatic and lead to small temperature changes. Heat capacity is much smaller in a gas cavern, temperature evolutions are much faster, and heat transfer from the rock mass must be taken into account. Evolutions can be described through a relatively simple equation. A simple model describing pressure and temperature in the wellbore during gas withdrawal is obtained.

Key words: Salt caverns; Gas storage caverns; CAES; Thermodynamics of salt caverns; Water vapor in a salt cavern

Highlights

- Conditions leading to natural convection in a cavern
- Inversion of temperature gradient in a gas cavern
- Equations for evolution of stored fluid temperature during injection-withdrawal
- Adiabatic and non-adiabatic behavior of a liquid-storage cavern

- Thermodynamic behavior of a gas-storage cavern
- A closed-form solution for gas flow in a wellbore

Introduction

Salt caverns are deep cavities that are created by solution mining in both bedded and domal salt formations. They are connected to the ground level through a cased and cemented well. Strings are set in the well to allow injection or withdrawal of fluids into or from the cavern. These caverns range in volume from $V_c = 5000 \text{ m}^3$ to $5,000,000 \text{ m}^3$ and in depth from $H = 200 \text{ m}$ to $H = 2000 \text{ m}$. When solution mining is completed, the cavern can be filled with crude oil, LPG, natural gas, hydrogen, compressed air, etc. During operation, fluids are injected or withdrawn from the cavern, and heat is exchanged with the rock mass, leading to pressure and temperature changes in the cavern. An abundant literature — mentioned in the course of the paper — was dedicated to various aspects of these thermodynamic effects in specific sites. The objective of this paper is to describe some of these effects and to prove that their main features can be captured by relatively simple models.

1. Natural convection in a salt cavern

In this section, the theoretical conditions leading to the onset of natural convection in a cavern are discussed.

1.1. Fluid pressure in a cavern

In a gas-storage cavern, gas pressure (P) at casing-shoe depth (H) varies between a maximum and a minimum pressure, typically $P_{\min} = \gamma_{\min} H$, $P_{\max} = \gamma_{\max} H$, where $\gamma_{\min} = 0.6 \times 10^{-2} \text{ MPa/m}$, and $\gamma_{\max} = 1.8 \times 10^{-2} \text{ MPa/m}$ are the minimum and maximum “gradients”. Maximum gas pressure is smaller than geostatic pressure, $P_{\infty} = \gamma_R H$, $\gamma_R = 2.2 \times 10^{-2} \text{ MPa/m}$. where γ_R is the average volumetric weight of the rock formation. In natural gas caverns, gas used to be injected in summer and withdrawn in winter; at a $H = 1300\text{-m}$ depth, for instance, pressure varied from 8 MPa to 22 MPa, typically. Prompted by the needs of energy traders, cycles tend to be more frequent. In compressed air storage,

pressure is cycled every day between 5 MPa and 7 MPa, typically. A gas-storage cavern contains gas and brine (hence, water vapor) as, during the initial gas fill (debrining) some brine, trapped into the sump at the cavern bottom, cannot be removed from the cavern. This brine, which is cold (see Section 4.2), plays a significant role in cavern thermodynamics.

In liquefied or liquid (or supercritical) products storage, a central tube is left in the wellbore, and brine is injected in the central tube when hydrocarbon is withdrawn from the annulus between the cemented casing and the central tube (and vice versa): most often, cavern pressure is halmostatic, or $P_h = \gamma_b H$, $\gamma_b = 1.2 \times 10^{-2} \text{ MPa/m}$ is brine volumetric weight. However, cavern pressure is higher than halmostatic during tightness tests and lower during workovers. In the following, cavern pressure (P_c) is the fluid pressure computed at casing-shoe depth.

1.2. Temperature in a cavern

1.2.1. Brine-filled cavern

Figure 1 displays the temperature profile along a vertical axis in the cavern and access well of a brine cavern at Etrez, France, operated by Storengy (¹Brouard, 1998). For this cavern, $V_c = 8\,000 \text{ m}^3$, and the cavern's top and bottom depths are 920 m and 960 m, respectively. On the left hand-side, a slope discontinuity can be observed at a 700-m depth. Its origin is in the contrast between the thermal conductivities of salt and marls. Heat flux through conduction is continuous through any discontinuity, $q = K_R dT_R / dz = K_{marl} dT_{marl} / dz$. For rock salt, $K_R = 6 \text{ W/m/}^\circ\text{C}$; for marls, $K_{marl} = 3 \text{ W/m/}^\circ\text{C}$. The thermal gradient dT / dz is divided by 2 at the interface between salt and marls. Another slope discontinuity, originating in convection, can be observed at shallow depth (a few dozen meters below ground level). The test is performed in February. At night, the wellhead is cold. Cold brine flows down in the wellbore to be replaced by warmer brine from below which, in turn, cools when in contact with the wellhead. The natural geothermal gradient is $G_R^\infty = dT_R^\infty / dz = 1.6 \times 10^{-2} \text{ }^\circ\text{C/m}$ in the salt formation. The thermal conductivity of the brine is $K_b = 0.6 \text{ W / m / }^\circ\text{C}$: if heat transfer in the cavern brine were through conduction, one would expect that the temperature gradient be larger in the brine mass than in the rock mass. In fact, the temperature gradient in the cavern is as small as $G_c = 3.4 \times 10^{-3} \text{ }^\circ\text{C/m}$, clear evidence of the effects of natural convection. At the cavern bottom, the brine is warmer (hence, lighter) than it is at the cavern top. Due to

buoyancy effects, brine rises; it cools along its upward path to the cavern top and ultimately returns to the cavern bottom, leading to the creation of convection cells. Brine is stirred by convection, and its temperature homogenizes. Numerical computations (²Karimi-Jafari 2007) using Boussinesq' linearization of the equations describing natural convection in the cavern were performed. Axisymmetry is assumed. The ADINA CFD modulus was used (k-ε model, no slip at cavern walls, velocity profile at cavern wall according to the Reichard law). On Figure 2, steady state flow is reached (after $1.1 \cdot 10^8$ s, or 12 days). Iso-values of the stream function $\psi = \psi(r, \theta)$, $v_r = -\partial\psi / \partial\theta$, $v_\theta = \partial\psi / \partial r$ are represented on Figure 2, left: three convection cells develop, the lower cell rolls counter-clockwise. On Figure 2, right, velocity magnitude is represented; it is largest along the axis of symmetry where it is faster than 0.2 mm/s (17 m/day).

Figure 3 displays the case of a brine-filled cavern at Stassfurt, Germany. The cavern volume is $V_c = 13\,000 \text{ m}^3$, and the cavern's top and bottom are 440 m and 650 m, respectively. As in the previous example, the thermal gradient in the cavern is much smaller than the natural geothermal gradient, which is $G_R^\infty = 1.8 \times 10^{-2} \text{ }^\circ\text{C/m}$ in the salt formation — clear evidence of the onset of convection.

1.2.2. Gas-filled cavern

In a gas cavern, convection also can be observed. In fact, onset of convection is certain in a brine cavern (see Section 1.3). This is less true in a gas cavern and, in many cases, only the upper part of the cavern is the seat of perennial natural convection, as explained below.

Figure 4 represents the temperature profile in the S107 natural gas cavern at Stassfurt, Germany. ⁴Klafki et al. (2003) mention that the top and bottom of the cavern are 968 m and 1070 m, respectively. However, Fig. 4 suggests that actual depths might be slightly larger. Cavern volume is $V_c = 288\,000 \text{ m}^3$; it can be divided in three parts, whose radii are, from top to bottom, 20 m, 60-65 m and 46 m. The cavern had been kept at rest for 2.5 months before the temperature log was run. Thermal convection is active in the upper part of the cavern: the thermal gradient there is much smaller than the geothermal gradient in the rock mass ($G_R^\infty = 1.35 \times 10^{-2} \text{ }^\circ\text{C/m}$). In the lower part of the cavern, the gradient sign changes (gradient *inversion*) slightly above the cavern bottom. The lower part of the logging tool was below the

brine-gas interface when the temperature at the lowest point on Fig. 4 was measured, which means that this point is representative of the temperature at the interface depth. Gradient inversion will be discussed in Section 4.

1.3. First condition for onset of natural convection

Gas thermal constants are given in Table 1 (at $\theta=25^{\circ}\text{C}$ and $P_N = 0.10325$ MPa). Heat capacity, C_p , is an increasing function of pressure. At 20 MPa and 30 °C, it is 3400 J/kg/°C for CH4 and 1600 J/kg/°C for air (Schlichtenmayer et al. 2015). In this paper, we are mainly interested in a semi-quantitative description of the phenomena observed in actual caverns – rather than in accurate numerical computations. For this reason, dependency of the thermal parameters upon temperature is not taken into account. Note, however, that it can play a role. For instance, it is known that, for a given geometry, rotation of convection cells is in opposite directions when a gas and a liquid is considered, as the sign of $d\mu/dT$ (μ is the fluid viscosity) is not the same.

Mechanical equilibrium in a fluid requires that temperature be a function of depth only. At equilibrium, $\text{grad } P = \rho(P,T)\vec{g}$; hence, $\text{curl}(\rho\vec{g}) = \left(\frac{\partial\rho}{\partial T}\right)\vec{g} \wedge \text{grad } T = \vec{0}$ and $T=T(z)$. Such an equilibrium is stable when the temperature gradient is oriented upward ($G = dT/dz < 0$). A downward-oriented temperature gradient ($G = dT/dz > 0$) is not a sufficient condition for onset of natural convection (i.e., instability). Even in an infinite fluid medium (No viscous boundary layer at cavern wall or viscous dissipation in the fluid body is considered.), this gradient must be larger than a certain threshold, the adiabatic gradient, which can be defined as follows. When a fluid particle rises by $dz < 0$, its pressure immediately decreases by $dP = \rho g dz$ to reach mechanical equilibrium with the surrounding fluid. Because of this expansion, liquid temperature drops. As a consequence, heat transfer from the surrounding fluid takes place to reach thermal equilibrium. However, this process is much slower than the mechanical process; in fact, when particle rise is fast, insufficient time is given for heat transfer, and particle expansion is adiabatic. In the case of a non-viscous ideal gas, an adiabatic expansion can be described as $Tds = C_p dT_{ad} - dP/\rho = 0$, where s is the gas entropy per unit of mass, C_p is the gas thermal capacity (when gas pressure is kept

constant), $\rho = P / rT$ is its density, from which $dT_{ad} = g dz / C_p$. This temperature change must be compared to the geothermal temperature change, or $dT_c = G_R^\infty dz$. When $G_{ad} = g / C_p < G_R^\infty$, the rising particle is warmer (and lighter) than its environment and keeps rising: convection appears. Gravity acceleration is $g = 10 \text{ m/s}^2$; heat capacity is $C_p = 3400 \text{ J/kg/}^\circ\text{C}$ for natural gas (at 30°C and 20 MPa), $C_p = 1100 \text{ J/kg/}^\circ\text{C}$ for air (at 30°C and 7 MPa), and $C_p = 14\,000 \text{ J/kg/}^\circ\text{C}$ for hydrogen. In other words, as $G_R^\infty = 1.6\text{--}1.8 \times 10^{-2} \text{ }^\circ\text{C/m}$ in a salt formation, onset of convection is likely in an air-storage facility ($g / C_p = 9 \times 10^{-3} \text{ }^\circ\text{C/m}$), highly likely in a natural gas storage ($g / C_p = 3 \times 10^{-3} \text{ }^\circ\text{C/m}$) and certain in a hydrogen storage ($g / C_p = 0.7 \times 10^{-3} \text{ }^\circ\text{C/m}$). In the case of a liquid, $dT_{ad} = \alpha T dP / \rho C_p$; $\alpha = -\partial \rho(T, P) / \partial T$ is the thermal expansion coefficient (for a gas, $\alpha T = 1$.) For brine, $\alpha_b = 4.4 \times 10^{-4} / ^\circ\text{C}$; $T = 300 \text{ K}$ is typical, $C_p^b = 3800 \text{ J/kg/}^\circ\text{C}$, and $G_{ad} = 3.5 \times 10^{-4} \text{ }^\circ\text{C/m}$ is much smaller than the geothermal gradient: onset of convection is certain in a brine-filled cavern, as suggested by the examples shown on Figs. 1 and 3. However, these simple arguments fail to take into account several factors, which will be discussed below.

1.4. Induced geothermal gradient

In the previous section, it was assumed that the geothermal gradient is the same in the cavern (G_c^∞) as in the rock mass (G_R^∞). In fact, because gas (or liquid) conductivity is less than rock conductivity, the geothermal gradient at rest (when no convection takes place) is larger in the cavern fluid than in the rock mass. The adiabatic gradient (G_{ad}) must be compared to G_c^∞ rather than G_R^∞ . Let T_R and T_F be the temperature distributions at rest in the rock mass and in the cavern fluid, respectively. Both are harmonic functions, $\Delta T_R = 0$ and $\Delta T_F = 0$. Consider, for instance, the case of a spherical cavern, radius a , whose origin is at the center of the cavern whose depth is H ; spherical coordinates are used, and $z = H + r \cos \theta$ is oriented downward. Rock temperature and fluid temperature are $T_R(r, \theta) = T_\infty(H) + G_R^\infty r \cos \theta + A a^3 \cos \theta / r^2$ and $T_F = T_\infty(H) + G_c^\infty r \cos \theta$, respectively; A and G_c^∞ are two constants. At the cavern wall, temperatures and heat flux are continuous, $T_R = T_F$ and $K_F \partial T_F / \partial n = K_R \partial T_R / \partial n$, from which it is inferred that the geothermal gradient at rest in the cavern is $G_c^\infty / G_R^\infty = 3K_R / (2K_R + K_F)$. In most cases, K_F is much smaller than K_R (for rock

salt, $K_R = 6 \text{ W/m/}^\circ\text{C}$; for brine, $K_F = 0.6 \text{ W/m/}^\circ\text{C}$; for air or for hydrogen, $K_F = 0.026 \text{ W/m/}^\circ\text{C}$ or $0.185 \text{ W/m/}^\circ\text{C}$ respectively — see Table 1) and the “induced” geothermal gradient in the fluid at rest is larger than the geothermal gradient in the rock mass (for air, $G_c^\infty = 1.5 G_R^\infty$), making onset of convection somewhat easier than predicted in Section 1.3. The induced geothermal gradient also depends on cavern age. During the leaching period, which is several years long, cold water pumped from a lake or a shallow aquifer is injected in the cavern. When leaching is completed, the cavern brine is colder than the rock mass by a couple of dozen of $^\circ\text{C}$. After debrining, the cavern walls are cold (⁷Walden et al., 2001), and the same can be said of the brine left at the cavern bottom. Warming brine and rock mass is slow (⁸Fosse and Røvang, 1998), as discussed later. The transient induced geothermal gradient is difficult to compute.

1.5. Wet gas and dry gas

In many cases (Section 2), cavern gas is saturated with water vapor. (Equilibrium with sump brine must be reached.) In principle, when computing adiabatic gas expansion, the effect of condensation must be taken into account, and the wet adiabatic gradient is smaller than the dry adiabatic gradient (⁹Sivoukhine, 1982, p. 499-501). Let m_g , m_v , m_L be the mass of gas, water vapor and liquid water, respectively. When condensation starts, $\dot{m}_L = 0$ and $\dot{m}_L + \dot{m}_v = 0$. The vapor pressure at saturation, $P_v^s(T)$, is a function of temperature only. During adiabatic evolution, the entropy of the mixture, $s = \sum_i m_i s_i$, $i = v, L, g$, is constant, and

$$(s_v - s_L)dm_v + \frac{m_v}{T}(C_p^v dT_{ad} - \frac{dP_v^s}{\rho_v}) + \frac{m_g}{T}(C_p^g dT_{ad} - \frac{dP_g}{\rho_g}) = 0 \quad (1)$$

Let $L = (s_v - s_L)T$ be the latent heat capacity for the liquid-vapor transition; taking into account $dP_v^s / dT = L\rho_v / T$ (Clapeyron’s law), $dP_g = \rho_g g dz$ (mechanical equilibrium), and $P_i = \rho_i r_i T$ for $i = v, g$ (state equations)

$$\left. \frac{dT}{dz} \right|_{wet} = \frac{g}{C_p^g \left(1 + \frac{r_g P_v^s}{r_v P_g} \left[L \left(\frac{L}{r_v T^2} - \frac{2}{T} \right) + C_p^v \right] \right) / C_p^g} \quad (2)$$

The difference between the wet adiabatic gradient G_{ad}^{wet} and the dry adiabatic gradient G_{ad}^{dry} (which equals g / C_p^v) is proportional to the P_v^s / P_g ratio. This ratio between vapor pressure and gas pressure is not very small in the atmosphere ($P_g = P_{air} = 0.1$ MPa; at $\theta = 30^\circ\text{C}$, $P_v^s(30^\circ\text{C}) = 4.25$ kPa and $P_v^s / P_{air} = 4\%$, see Section 2) and the difference is significant when meteorological phenomena are concerned. The same is not true for a gas cavern, in which $P_g = 10$ MPa, and, when $\theta = 50^\circ\text{C}$, $P_v^s(50^\circ\text{C}) = 9$ kPa are typical.

1.6. Second condition for onset of natural convection

The condition for onset of natural convection, $G_{ad} < G_c^\infty$, is a necessary condition. In the analysis, viscous effects were not taken into account. In fact, in a cavity of (very) small dimensions, viscosity can impede natural convection. The set of equations describing natural convection in a cavern (conservations of mass, energy and momentum) must be solved numerically. However, steady-state convective flow can be described through linearized equations (Boussinesq approximation) (¹⁰Landau and Lifschitz, 1971). This set of equations highlights the significance of three dimensionless constants, Prandtl, Grashof and Nusselt numbers: $Pr = \nu/k$, where ν = kinematic viscosity, k = gas thermal diffusivity; $Gr = g\alpha Ga^4 / \nu^2$, where a = cavern characteristic length, α = the thermal expansion coefficient of the fluid, and $G = G_c - G_{ad}$, where G_c is the actual thermal gradient in the cavern; and $Nu = \bar{H}a / K_F$ where \bar{H} is the heat-transfer coefficient at the cavern wall (defined in Section 5.1). Natural convection starts when G is larger than the value given by a certain combination of Gr , Pr and Nu that depends on cavern shape and boundary conditions. For an elongated cylindrical cavern, convection appears when the Rayleigh number $Ra = Pr \cdot Gr$ is larger than a value that depends on Nu and is smaller than 10^3 (¹⁰Landau and Lifschitz, 1971). The Prandtl number can be written $Pr = \mu C_p / K$; for most gases (Table 1), dynamic viscosity is in the range $\mu = 1 - 2 \times 10^{-5}$ Pa.s $C_p = 1 - 2 \times 10^3$ J/kg/K, and $K = 2 \times 10^{-5}$ Pa.s. Also, $g = 10$ m/s², $G = 1.5 \times 10^{-2}$ °C/m, $\alpha_g = 3 \times 10^{-3}$ /°C and, for standard caverns, $a = 10$ -50 m.

The Rayleigh number ($= Pr.Gr$) is larger than 10^{10} , a figure for which natural convection is turbulent. In other words, a cavern contains so large a fluid body that the condition relative to the adiabatic gradient is a sufficient condition for onset of convection. For instance, ⁴Klafki et al., 2003 computed natural convection in the cavern represented on Fig. 4. A stable vortex (The horizontal component of gas speed was high.) was observed in the lower part of the cavern, an independent convection cell (upstream at the wall, downstream in the center) was observed in the main cavern, and flow was unsteady in the narrow upper part. Maximal velocities were smaller than 0.5 m/s. This computation highlights the importance of cavern shape on the pattern of convection cells pattern (see also Fig. 2).

2. Water vapor in the cavern

Moisture content is a critical issue when operating a gas cavern. It was proven in Section 1.5 that its influence on the onset of natural convection is minor. However, when gas containing water vapor is withdrawn from a cavern, gas pressure and temperature drop in the wellbore, leading to the possible formation of hydrates in the upper part of the wellbore (a severe operational problem, discussed in Section 6). In fact, natural gas injected in a cavern is relatively dry. In France, for instance, water vapor concentration (or $c_v = m/V_N = P_v P_N / r_v T_N P_g$, where the subscript N is for normal conditions, $P_N = 0.10325$ MPa, and $T_N = 273$ K.) must be less than $c_v = 46$ mg/Nm³ (¹¹Louvet et al., 2018). However, there is a thin brine layer at the cavern wall (70 cm³/m², see ¹²Kökritz and Walden, 1994) and a (much larger) amount of brine in the sump at the cavern bottom mentioned in Section 1.1. Equilibrium requires, at least in principle, that brine is saturated with dissolved natural gas and that cavern gas is saturated with water vapor. The partial pressure of water vapor at saturation, or P_v^s , is a function of temperature (for instance, Rankine's formula predicts: $P_v^s [T] = f P_{atm} \exp(13.7 - 5120/T)$, with T in K; for pure water, $f = 1$, and for saturated brine, $f = 0.75$, according to Raoult's law).

Following gas injection or withdrawal, temperature changes in the cavern, and gas is either over- or under-saturated with water vapor in the new temperature conditions. Water vaporizes or condenses until, after some time, equilibrium is reached again. Consider an air-storage cavern of volume $V_c = 400\,000$ m³: air is withdrawn rapidly from the cavern, whose

pressure drops from 7 MPa to 5 MPa: air temperature drops from $\theta = 50^\circ\text{C}$ to $\theta = 30^\circ\text{C}$ (see Section 5.12), and saturated vapor pressure drops from $P_v^s(50^\circ\text{C}) = 9 \text{ kPa}$ $P_v^s(30^\circ\text{C}) = 3.15 \text{ kPa}$. The mass of water vapor in air or $m_v = P_v^s V_c / r_v T_c$, where $r_v = 460 \text{ J/kg/}^\circ\text{C}$, drops from 24,000 kg to 9,000 kg. Water vapor concentration also drops, from 1050 mg/Nm^3 to 515 mg/Nm^3 . When withdrawal starts, gas becomes oversaturated with water vapor, and it hazes — or even rains — in the cavern. Latent heat generated by condensation of $\delta m = 15,000 \text{ kg}$ of water ($L\delta m = 3.75 \cdot 10^{10} \text{ Joules}$) is provided to the gas mass, and the gas temperature drop is slightly smaller than predicted from gas-expansion equations (-20°C), by $\delta\theta = L\delta m / \rho_g C_p^g V_c = 1^\circ\text{C}$, approximately ($\rho_g C_p^g = 0.8 \times 10^5 \text{ J/m}^3/^\circ\text{C}$). At the end of the withdrawal phase, air warms to return to thermal equilibrium with the rock mass ($\theta = 50^\circ\text{C}$). Later on, warm gas is injected in the cavern, pressure and temperature increase. Vaporization then takes place at the vicinity of the interface; latent heat now is provided by brine and gas close to the brine/gas interface (rather than by the whole gas mass) — see Fig. 5 (Thaule and Gentsch, 1994). Note the temperature gradient inversion at a depth of 1280 m, discussed below. This resaturation process seems to be relatively slow. Köckritz and Walden (1994) and Köckritz et al. (1996) studied vaporization rate at the brine-gas interface at the laboratory scale and proposed the following law: $\dot{m}_v = A\bar{\beta}(P_s - P_v(t)) / r_v T$, where $m_v = P_v V_c / r_v T$, i.e., $\dot{m}_v = A\bar{\beta}(m_s - m_v(t)) / V_c$; A is the brine-gas interface area; and $\bar{\beta}$ is a constant. Vapor mass is the time constant for the vaporization process, $t_c = V_c / A\bar{\beta}$. The ratio between cavern volume (V_c) and interface area (A) in an elongated cylindrical cavern typically is $V_c/A = 200 \text{ m}$. Köckritz et al. (1996) measured $\bar{\beta} = 5 \times 10^{-5} \text{ m/s}$ (at the lab); they suggest that this figure must be increased by a factor of 20 when convection is active, leading to $t_c = 130 \text{ days}$ (no convection) and $t_c = 6.5 \text{ days}$ (convection is active). Warm gas injection at the cavern top and slow re-saturation may explain why withdrawn gas often is under-saturated with vapor, see Section 4.2. In other words, a pressure cycle is not a symmetric process: it provides condensation heat to the entire gas body and withdraws vaporization heat from the vicinity of the brine/gas interface (In the example above, vaporization heat is $L\delta m = 3.75 \times 10^{10} \text{ J}$, an amount which is able to cool a $10,000 \text{ m}^3$ volume of brine by 1°C).

3. Examples of temperature evolution during gas injection-withdrawal

In this section, examples of temperature profile evolution when gas pressure varies significantly are presented and the role of vapor content discussed.

3.1. S107 Cavern, Stassfurt, Germany

Cavern S107 was described briefly in Section 1.2.2 and Fig. 4. During a 14-day period, three gas withdrawals (60,000 Nm³/hr over 20 hours, 100,000 Nm³/h over 12 hours, 142,000 Nm³/hr over 7 hours, followed by a 77-hour standstill) and a 61,000-Nm³/hr injection over 61 hours were managed (for natural gas, 1 Nm³ = 0.68 kg). Temperature profiles were measured through a Raman optical fiber (Fig. 6); they remain almost vertical during withdrawal, evidence of the perennial effect of thermal convection in the main cavern body. The height of the cool zone in the lower part of the cavern (below the gradient inversion depth) increases. A separate convection cell is present in the upper (and thinner) part of the cavern, where gas cooling is slower. Gas temperature immediately above the gas/brine interface (noted by the sinker bar being in the brine, ⁴Klafki et al., 2003) cools significantly, and it is reasonable to assume that brine below the interface cools accordingly.

3.2. Huntorf caverns

Figure 7 (¹⁵Quast, 1983) displays temperature profiles in the NK1 cavern of the Huntorf compressed-air storage facility not long after the first air fill. (The shape of the Huntorf caverns is represented on Fig. 14). At the cavern bottom, the sump brine is still quite cold. In Fig. 7, air flowrates (in kg/s) are positive when gas is withdrawn. Each of the four profiles corresponds to a given instant of an injection-withdrawal sequence (1. standstill; 2. end of withdrawal; 3. end of a standstill and 4. end of an injection phase). The temperature vertical gradient is inverse (Its sign is opposite to the geothermal gradient — i.e., no convection takes place), and its value is $G_c = dT / dz = -10 \times 10^{-3} \text{C/m}$ (z oriented downward). According to ¹⁵Quast (1983), this inverse gradient can be explained by brine temperature at the cavern bottom, which is much colder than the rock-mass temperature. This gradient remains constant even when the average gas temperature changes significantly. This explanation is reasonable; however, it must be noted that, in addition, onset of convection is more difficult

in the case of air (or oxygen). More recent profiles would have shed some light on this, but, unfortunately, they are not available.

3.3. Enterprise cavern

Figure 8 (¹⁶Skaug et al., 2010) displays temperature profiles before and after a natural gas withdrawal in a salt cavern operated in Texas by Enterprise. Measurements were performed in April and May 2009, before and after movement of gas. (Of the initial inventory, 22% was withdrawn, and the initial cavern pressure at casing shoe depth was $P = 1.85 \times 10^{-2} H$, P_c in MPa, H in m.) When the cavern was leached out is unknown; however, it is known that a sonar survey had been run in 2002, 7 years before the 2009 test. The cavern volume is $V_c = 430\,000\text{ m}^3$ (2.6×10^6 bbls), and its depth is between $H = 3600\text{ ft}$ (1100 m) and 3960 ft (1200 m). Here, again, thermal convection is impeded in the lower part of the cavern; at rest, gradient inversion depth is 3720 ft. Figure 8 suggests that temperature also was measured below the brine-gas interface. More than 7 years after debrining was completed, brine temperature is significantly colder (32 °C) than that of the cavern gas — and colder still after gas withdrawal (27 °C). It is likely that, in absence of brine, the gas temperature should have dropped by a larger amount in the lower part of the cavern. Brine convection takes place below the brine-gas interface, effectively cooling at least the upper part of the brine body. However, brine temperature remains warmer than gas temperature in the lower part of the cavern, and thermal convection also is triggered in the gas, leading to an almost vertical temperature profile in the lower part of the cavern. This tentative explanation (Few data are available.) is discussed in more detail in Section 4.

4. Inversion of temperature gradient and tentative synthesis

4.1. Temperature distribution at the cavern bottom

It was noted in the above examples (Figs. 4 to 6 and Fig. 8) that the temperature gradient, which was oriented upward in the main cavern body, experiences an inversion in the lower part of the cavern. This seems to be a general property of gas-storage caverns: "... frequently, measurements in caverns reveal a drop in temperature a few meters above the brine level. This is attributable to the consumption of heat energy for evaporation" (¹⁷Berger et al., 2002, p.56); "The brine remaining in the bottom of the cavern is usually cooler than

the stored gas” (¹⁸Istvan, 1996, p. 3); and “Above the gas-brine level the temperature decreases significantly which is a typical effect observed in caverns” (¹⁹Schneider et al., 2002, p. 263). Temperature profiles reproduced by ²⁰Karimi-Jafari et al. (2014), ²¹Krieter and Gotthardt (2015), ²²Osnes et al. 2007 or ¹⁷Berger et al., 2002 (see Fig. 9) display the same pattern: along the cavern’s axis of symmetry, the thermal gradient is positive and low in the upper part of the cavern ($dT/dz = 6 \times 10^{-3} \text{C/m}$); it is negative and high above the sump at the cavern bottom ($dT/dz = -0.15 \text{C/m}$). The ratio between cold-zone height and cavern height (h/h) is defined on Fig. 9; it varies from 10% (on Figs. 4 and 5) to 60% (Kiel 102 cavern, discussed by ¹⁹Schneider et al., 2007) and 100% (at Huntorf, see Fig. 7). It was said that, after the first gas injection (“debrining”), some brine (1000-10,000 m³) is left at cavern bottom in the sump formed by the insoluble fractions that fell during solution mining. This brine is cold and remains cold for a while. It seems reasonable to infer that the lower part of the cavern gas is the seat of thermal conduction and water vapor diffusion (two slow transfer processes) and that the upper part is the seat of natural convection, leading to much faster vapor transfer. Temperature gradient (and heat flux) seem to be discontinuous at the interface between these two parts of the cavern; in fact, heat is transferred by conduction from a thin interface to both the lower and upper zones and by conduction + convection in the upper zone. Let S_0 be the surface on which convection vanishes, $\vec{u} = \vec{0}$, where \vec{u} is the gas speed, and S_1 is a surface slightly above S_0 . When density changes are neglected, $\int_{S_1} \vec{u} \cdot \vec{n}_1 dA_1 = 0$. The energy equation is written $d(\rho C_p^g T)/dt = \text{div}(-K_g \overline{\text{grad}T})$; when steady state is considered, heat is withdrawn from the gas slice between S_0 and S_1 through upward and downward conduction, $\int_{S_0} -K_g \partial T / \partial n_0 dA_0 + \int_{S_1} -K_g \partial T / \partial n_1 dA_1$, and brought into the gas slice through convection, $\int_{S_1} \rho_g C_p^g T (\vec{u} \cdot \vec{n}_1) dA_1$.

4.2. A tentative synthesis

Main findings of Sections 1 to 4.1 can be summarized as follows. The theoretical conditions for onset of natural convection in a cavern are met: the Rayleigh number, $Ra = Pr \cdot Gr$, is so large that convection, when it exists, is turbulent; the adiabatic gradient is significantly smaller than the geothermal gradient (Although this is less true for air storage.); and in a gas cavern, the wet adiabatic gradient almost equals the dry gradient. Convection stirs gas

cavern, whose temperature and moisture content should be homogeneous. In fact, in actual caverns, the general picture is more complex. Natural convection is always present in a liquid-filled cavern. In a gas-filled cavern, the cavern can be subdivided in two parts: (1) the upper part is the seat of natural convection, the temperature gradient is oriented upward, heat and moisture transfer is fast; and (2) in the lower part, the temperature gradient is oriented downward, no convection takes place, and heat transfer is slow. The gradient inversion depth varies widely from one cavern to another. One exception is the Huntorf storage cavern, where, at least early on, the temperature gradient was oriented downward in the entire cavern. In addition, it is observed that natural gas is not always vapor-saturated, especially at the cavern top (¹⁴Köckritz et al., 1994, ¹⁵Quast 1983), which can be related to the absence of convection immediately above the brine-gas interface where, to a certain extent, water vapor is trapped by the downward-oriented temperature gradient. These features can be related to the thermal history of a cavern and the presence of a cold sump at the cavern bottom. During the leaching phase, the rock mass is cooled significantly by circulation of cold water; several years after leaching is completed, the temperature at the cavern wall is significantly different from the geothermal temperature (⁷Walden et al., 2001). This is also true when brine at the cavern bottom, trapped in the insoluble sump, is considered: it is colder than the rock mass and conditions for onset of convection are not met in the lower part of the cavern. In addition, observations suggest that sump brine might remain colder than the rock mass during the entire cavern life because of several mechanisms. Heat is transferred from sump brine through the cavern walls and the brine-gas interface by conduction, a slow process — especially when gas is concerned. Gas is injected at the cavern top; it is warmer (60 °C) than cavern gas and under-saturated with water vapor; it may remain stagnant at the cavern top. Conversely, ²³Reitze and von Tryller (2005) suggest that, during gas withdrawal, a cold column forms at the center of the cavern that is colder than the gas that is closer to the cavern walls; after some time, this column drops to the cavern bottom. Both contribute to less effective convection. In Section 1.3, another reason was suggested. During a pressure cycle, brine at the bottom of the cavern is cooled during each cycle by two mechanisms: possible natural convection in the sump brine following a severe gas pressure drop (Enterprise cavern, Section 3.3); and, probably more frequent, vaporization following an increase in gas temperature during the warming phase of the cycle — two effects that have no counterpart during the cooling phase of the cycle.

This might explain why brine, which is cooler than the rock mass when cavern leaching is completed, remains consistently cold, at least when the cavern is submitted to frequent and severe pressure and temperature cycles. These explanations are tentative. Published temperature profiles are scarce and, in many cases, relevant information (cavern age, cavern pressure history before temperature is recorded) is missing.

5. Thermodynamic behavior of the stored fluid during injection-withdrawal

5.1. General equations

In this section, it is accepted that, when analyzing the overall thermodynamic behavior of a cavern, it is sufficient to consider the cavern's average temperature (T_c) and pressure (P_c). The set of equations satisfied by the rock mass and the stored fluid can be written

$$\frac{\partial T_R}{\partial t} = k_R \Delta T_R \quad (3)$$

$$G_R^\infty(z) = dT_R^\infty(z) / dz \quad T_R(t=0) = T_R^\infty(z) \quad (4)$$

$$T_R = T_c \quad \text{at cavern wall} \quad (5)$$

$$v_c = v_c(P_c, T_c) \quad h_c = h_c(P_c, T_c) \quad (6)$$

$$m(\dot{h}_c - v_c \dot{P}_c) = Q + \langle \dot{m} \rangle (h_{inj} - h_c) - L\dot{m}_v \quad (7)$$

$$d(\rho_f V) / dt = \dot{m} \quad (8)$$

$$\dot{V}_c / V_c = \beta_c \dot{P} - \alpha_R \Phi[T(\underline{x})] + \dot{\epsilon}^{cr} \quad (9)$$

Equations (3, 4, 5) describe heat transfer through the rock mass: Equation (3) is the heat-conduction equation in the rock mass (k_R is the thermal diffusivity of salt) ; Eq. (4) contains the boundary and initial thermal conditions, respectively (Other initial conditions also can be considered); and Eq. (5) is the boundary condition at the cavern wall. In fact, many authors (e.g., ²⁴Kushnir et al., 2012), assume that fluid temperature at the cavern wall is not equal to the rock temperature, and some surface heat-transfer coefficient (\bar{H}) must be taken into account, $\bar{H}(T_R - T_c) = K_R \partial T_R / \partial n$. Selecting the value of this parameter (\bar{H}) is difficult. (There

is a lack of field data, and this value depends strongly on roughness of the cavern wall and convection intensity.) In fact, the turbulent boundary layer at the cavern wall is likely to be thin, making \bar{H} quite small; in this paper, it is assumed that temperature is continuous at cavern wall. The stored fluid can be characterized [see Eq. (6)] by its state equation — i.e., the specific volume, $v_c = 1/\rho_c$, of the stored fluid as a function of P_c and T_c , $v_c = v_c(P_c, T_c)$, and by a thermodynamic potential — for instance, its enthalpy per unit of mass, $h_c = h_c(P_c, T_c)$. The energy balance equation is shown in Eq. (7), where $Q = \int_{\Sigma} K_R \partial T_R / \partial n d\Sigma$ is the heat flux transferred from the rock mass to the gas through the cavern wall, and m is the mass of fluid in the cavern. Kinetic energy is disregarded: when gas is withdrawn/injected from/to the cavern, the gas rate in the wellbore, which has a cross-sectional area smaller than 0.1 m², typically is $u = 10$ m/s, (less in the case of a liquid). Rates are much smaller when considering the cavern's bulk volume, whose cross-sectional area typically is several thousands of m². The system is open, and mass is exchanged with the wellbore through the cavern chimney; $\langle \dot{m} \rangle = \dot{m}$ when $\dot{m} > 0$ and $\langle \dot{m} \rangle = 0$ when $\dot{m} < 0$. At the cavern entrance, fluid pressure is continuous. When gas is injected, temperature is not continuous, and a jump in enthalpy must be considered, $h_{inj} - h_c = C_p(T_{inj} - T_c)$, where T_{inj} is the temperature of the injected fluid. No such term exists when fluid is withdrawn ($\dot{m} < 0$). In a gas cavern, gas often is compressed before being injected; however, in most cases, T_{inj} is smaller than 60 °C to prevent excessive thermal expansion of the wellhead and steel casings. The latent heat associated with condensation/vaporization is $L\dot{m}_v$; it will not be considered in the following. Equation (8) is the mass-conservation equation for the stored fluid, V is the fluid volume; Eq. (9) describes cavern volume evolution, where V_c is the cavern volume, $V = V_c$, when there is only one fluid in the cavern. The elastic compressibility of the cavern is β_c , which depends on the elastic properties of the rock mass (\bar{E} and $\bar{\nu}$) and cavern shape — not on cavern size. For an idealized spherical or cylindrical cavern, $\beta_c = 3(1+\bar{\nu})/2\bar{E}$ and $\beta_c = 2(1-\bar{\nu}^2)/\bar{E}$, respectively; $\beta_c = 1.4 \times 10^{-4}$ /MPa is typical, larger values are found in flat caverns (²⁵Bérest et al., 1999); $\dot{\epsilon}^{cr}$ is the cavern creep-closure rate, which depends on cavern pressure history, cavern depth and shape, and the creep properties of salt. Generally speaking, $\dot{\epsilon}^{cr} < 0$, except after a swift pressure increase (²⁶Bérest et al., 2015), and $\dot{\epsilon}^{cr}$ varies from $\dot{\epsilon}^{cr} = -10^{-5}$ /yr at a 250-m depth to $\dot{\epsilon}^{cr} = -$ several %/yr at a 2000-m depth. Term $\alpha_R \Phi[T(x)]$ is

the thermal expansion of the cavern when the rock mass is warmed or cooled. It is zero in the case of an idealized spherical or cylindrical cavern and, in general, is exceedingly small (²⁷Karimi-Jafari et al. 2007); it is disregarded in the following, $\Phi = 0$, and $\alpha_r = 4 \times 10^{-5} / ^\circ\text{C}$ is the thermal expansion coefficient of rock salt. One may ask whether cavern volume, or V_c , which experiences large changes in the long term, must be updated. In the following, focus is put on two problems: effects of a fast gas pressure drop (Sections 5.5 to 5.12), during which cavern additional shrinkage is small; and long-term evolution (centuries) of a brine-filled abandoned cavern, for which extremely accurate predictions are out of reach. For this reason, cavern volume is not updated; it could be; an interesting attempt can be found in ²⁸Kalhedi et al.

5.2. Temperature evolution in a liquid-filled cavern

In the case of a liquid, $\dot{v}_L = v_L^0(\alpha_L \dot{T}_c - \beta_L \dot{P}_c)$ and $\dot{h}_L = C_p^L \dot{T}_c - \beta_L v_L^0 P \dot{P}$; Eqs. (7,8,9) can be rewritten:

$$m(C_p^L \dot{T}_c - \alpha_L T_c v_L^0 \dot{P}) = Q + \langle \dot{m} \rangle C_p^L (T^{inj} - T_c) \quad (10)$$

$$\dot{m} / m = -\alpha_L \dot{T}_c + \beta \dot{P}_c + \varepsilon^{cr} \quad (11)$$

where, in principle, $\beta = \beta_c + \beta_b$; β is the cavern compressibility factor. When the cavern is filled with brine, $\beta = 4 - 5 \times 10^{-4} / \text{MPa}$ is typical (²⁵Bérest et al., 1999). Note, however, that when the fluid is saturated with brine, mass exchange takes place between the rock mass and the cavern brine. Concentration at saturation is an increasing function of temperature and pressure. For instance, according to Le Chatelier-Braun principle, when pressure increases, dissolution occurs, and the volume of formed brine is smaller than the sum of the initial volume of brine and the volume of dissolved salt. This effect is taken into account by increasing the value of the coefficient β_b slightly (see ²⁵Bérest et al., 1999). This simplification holds when long-term evolutions are considered — in other words, when enough time is left to reach saturation.

5.3. Temperature evolution in an open liquid-filled cavern

In an open brine-filled cavern, when long-term evolutions are considered, $\dot{P}=0$ and $\dot{m}<0$. (In most cases, brine is expelled from the cavern; see, for instance, ²⁹Brouard et al., 2013). Thus, Eq. (10) can be rewritten:

$$\rho_L V_c C_P^L \dot{T} = \int_{\Sigma} K_R \frac{\partial T_R}{\partial n} d\Sigma \quad (12)$$

an equation which, together with Eqs. (3,4,5), allows computation of cavern temperature evolution. Let a be a characteristic length of the cavern. Taking into account $K_R = \rho_R C_P^R k_R$, two characteristic times can be built, a^2 / k_R and $a^2 / k_R \chi$, where $\chi = \rho_R C_P^R / \rho_L C_P^L$ is the ratio between the volumetric heat capacities of rock and fluid. In the case of a liquid (brine or oil), these two quantities are of the same order of magnitude (for brine, $\chi = 0.55$), and any of them can be considered to discuss orders of magnitude. Figure 10 displays brine temperature evolution after leaching is completed in a cavern with volume $V_c = 8,000 \text{ m}^3$ cavern, and an average depth $H = 950 \text{ m}$. A typical characteristic length is $a = V_c^{1/3} / 2$, and the characteristic time is $V_c^{2/3} / 4k_R = 1 \text{ year}$, a period after which 75% of the initial temperature gap vanishes (Fig. 10). Even in such a small brine-filled cavern, brine warming is a slow process. In a larger cavern (for instance, one with a volume of $V_c = 512,000 \text{ m}^3$), the characteristic time is 40 years, and brine warming is very slow. A more general discussion, including the effects of cavern shape, can be found in ²⁷Karimi-Jafari et al., 2007.

5.4. Pressure and temperature evolutions in a liquid-filled cavern

With the cavern closed at ground level, $\dot{m}=0$ and $-\alpha_L \dot{T}_c + \beta_L \dot{P}_c + \dot{\epsilon}^{cr} = 0$; Eq. (10) is written

$$\left(C_P^L - \frac{\alpha_L^2 T_c}{\beta \rho_L} \right) \dot{T}_c + \frac{\alpha_L T_c}{\beta \rho_L} \dot{\epsilon}^{cr} = \frac{1}{m} \int_{\Sigma} K_R \frac{\partial T_R}{\partial n} d\Sigma \quad (13)$$

where $\rho_L = 1/v_0$, $C_P^L - \alpha_L^2 T / \beta \rho_L = C_v^L$ is not very different from $C_v^L = C_P^L - \alpha_L^2 T / \beta_L \rho_L$. For brine, $\beta = 4 \times 10^{-4} / \text{MPa}$, $\alpha_b = 4.4 \times 10^{-4} / ^\circ\text{C}$, $T = 300 \text{ K}$, $\rho_b = 1200 \text{ kg/m}^3$, and $\alpha_L^2 T / \beta \rho_L = 120 \text{ J/kg/}^\circ\text{C}$ is significantly smaller than $C_P^b = 3600 \text{ J/kg/K}$. (Generally speaking, theoretical considerations lead to distinguishing between such notions as C_P^L , C_v^L , etc. It must be recognized, that, from a practical perspective, the difference is small in the case of a liquid.) In a $H = 1000\text{-m}$ deep cavern, when cavern pressure is halmostatic ($P_c = \rho_b g H$),

$\dot{\epsilon}^{cr} = -3 \times 10^{-4} / \text{yr}$ is typical (¹Brouard, 1998), and the effect of creep must be taken into account when \dot{T} is not much larger than $-\alpha_L T_c \dot{\epsilon}^{cr} / \beta \rho_L C_p^L = 0.05 \text{ }^\circ\text{C}/\text{yr}$. Now, when cavern temperature (rather than cavern pressure) is eliminated between Eqs. (10) and (11),

$$\beta' \dot{P}_c = -\dot{\epsilon}^{cr} + \frac{\alpha_L \chi}{V_c} \int_{\Sigma} k_R \frac{\partial T_R}{\partial n} d\Sigma \quad (14)$$

(β' is slightly distinct from β). This equation proves that in a shut-in cavern, in which, initially, the stored fluid is colder than the rock mass, cavern pressure increases due to two distinct phenomena: cavern creep closure, and brine expansion. The relative significance of these two phenomena depends on many factors, among which are cavern depth and cavern age (²⁶Karimi-Jafari et al., 2007). Cavern size plays no role in creep closure; however, the brine expansion rate is inversely proportional to the square of characteristic cavern length (a). The following example was made available by Storengy (³⁰Van Sambeek et al., 2005). Three caverns (A, B, C) were being leached in the same salt formation at comparable depth (Casing-shoe depths were 1450 m, 1465 m and 1590 m, respectively — Fig. 11). Hence, the cavern creep-closure rate ($\dot{\epsilon}^{cr}$) is almost the same in the three caverns. The three caverns were shut in, and the wellhead-pressure evolution was measured. Pressure increase rates were: $\dot{P} = 4 \text{ MPa}/\text{yr}$ in Cavern A ($V_c = 346,000 \text{ m}^3$), $5.9 \text{ MPa}/\text{yr}$ in Cavern B ($V_c = 147,000 \text{ m}^3$) and $10 \text{ MPa}/\text{yr}$ in Cavern C ($V_c = 48,600 \text{ m}^3$) — i.e., inversely proportional to $V_c^{2/3}$, which proves that brine thermal expansion was the main factor leading to pressure increase. As will be seen in Section 5, temperature increase is much faster in a gas-filled cavern.

These effects are important in an abandoned cavern. When thermal equilibrium is reached before abandonment, cavern pressure (P_c) increases due to creep; however, the creep rate decreases, as it is a non-linear function of the difference $P_\infty - P_c$ between geostatic pressure and cavern pressure. Ultimately, in principle, cavern pressure reaches geostatic pressure, and cavern creep closure extinguishes itself. In fact, ultimate pressure often is smaller than geostatic pressure because salt permeability, even if exceedingly small, allows some pressure relief. When thermal equilibrium is not reached before abandonment, thermal expansion leads to additional pressure increase, and geostatic pressure can be exceeded, with a risk of hydro-fracking. Abundant literature has been dedicated to this issue^{3,31,32,33}.

5.5. Fast injection or withdrawal in a liquid-filled cavern

In this paragraph, liquid is rapidly (less than one day) injected/withdrawn in/from a liquid-filled cavern. During such a short period, liquid evolution can be considered as adiabatic, $Q = 0$ (from Section 5.3); in addition, creep closure can be disregarded, $\dot{\epsilon}^{cr} = 0$. The case of an oil-filled cavern is discussed here. From Eqs. (10) and (11):

$$\dot{m} / m = -\alpha_o \dot{T}_c + \beta_o \dot{P}_c \quad C_p^o \dot{T}_c - \alpha_o v_0^o T_c \dot{P}_c = \frac{\langle \dot{m} \rangle}{m} C_p^o (T^{inj} - T_c) \quad (15)$$

The injection case ($\dot{m} > 0$) and the withdrawal case ($\dot{m} < 0$) are slightly different. In the withdrawal case ($\dot{m} < 0$), a simple formula is obtained:

$$\left. \frac{dT_c}{dP_c} \right|_{ad} = \frac{\alpha_o T_c}{\rho_o C_p^o} \quad (16)$$

In other words, a pressure drop generates a (small) temperature drop. This is illustrated clearly by tests performed in the *J* cavern of the Manosque site (in France) operated by Geostock (³³Gatelier et al., 2008) from June 21, 2005 to February 8, 2006. The cavern contained oil (and a small amount of brine at the cavern bottom) and had a volume of $V_c = 314\,000 \text{ m}^3$ and an oil density of $\rho_o = 800 - 840 \text{ kg/m}^3$. Temperature was measured at a 780-m depth, where oil pressure was 6.2 MPa.

Injections followed by withdrawals generated significant pressure changes (Fig. 12, left). Temperatures changed accordingly (Fig. 12, right). The ratio between measured temperature and pressure changes is $\dot{T}_c / \dot{P}_c = 0.16 \text{ }^\circ\text{C/MPa}$. Values $\alpha_o = 9 \pm 1 \times 10^{-4} / ^\circ\text{C}$; $T_c = 304 \text{ K}$ and $\rho_o C_p^o = 1.8 \pm 0.2 \times 10^6 \text{ J/m}^3 \cdot ^\circ\text{C}$ lead to a predicted value of $\alpha_o T_c / \rho_o C_p^o \approx 0.15 \text{ }^\circ\text{C/MPa}$, which is in good agreement with the measured value. Note that in a brine-filled cavern, the ratio $\alpha_b T_c / \rho_b C_p^b = 0.03 \text{ }^\circ\text{C/MPa}$ is five times smaller than in an oil-filled cavern, and temperature changes are more difficult to measure. However, they allow the resolution of temperature gages to be checked (²⁷Brouard et al., 2013).

5.6. Gas caverns

For an ideal gas, $Pv = r_g T$ and $\dot{h} = C_p \dot{T}$. (The value of $r_g = R / M_g$ depends on the gas considered.) In the following, we consider only swift (lasting a few days) gas injections/withdrawals. Creep closure can be neglected, $\dot{\epsilon}^{cr} = 0$; generally, gases are much more compressible than liquids, and cavern volume changes are neglected, $\dot{m}v + m\dot{v} = 0$. In a liquid-filled cavern, over a short period of time, heat flow from the rock mass can be neglected, $Q = 0$ (see Section 5.5). However, volumetric capacity of gas is much smaller than rock's (The χ -ratio is very small, see Section 5.3), and the significance of heat transfer in a gas-filled cavern must be discussed.

5.7. An example

An example is provided by ³⁴Heath and Benefield; a cycle lasting 6 months was applied in a cavern; pressure dropped from 16.2 MPa to 5.6 MPa, then returned to 16.2 MPa. The dotted line on Fig. 13 shows the computed adiabatic evolution. It is clear that neither the adiabatic nor the isothermal assumptions (see Figs. 6, 7, 8) fit the actual behavior of a gas-storage cavern; heat transfer from the rock mass must be taken into account.

5.8. Energy equation in a gas cavern

When cavern shape is known, precise computation of the heat flux from the rock mass (Q) can be done through numerical computations. However, when rapid injections/withdrawals are considered, simplifications can be made. Consider a δt -long rapid gas withdrawal — for instance, $\delta t = 4$ days. Pressure and temperature both drop. During such a short period, temperature changes are not given enough time to penetrate into the rock mass much deeper than $d = \sqrt{k_R t}$, where $k_R = 3 \times 10^{-6} \text{ m}^2/\text{s}$ is the thermal diffusivity of salt, or $d \approx 1\text{m}$. From the perspective of thermal conduction, cavern walls, as noted by ³⁶Crotogino et al., 2001 and ³⁷Krieter, 2011, can be considered as the sum, Σ , of small flat surfaces (a simplified description of small surfaces whose radius of curvature is not smaller than d .) Σ is the effective area of cavern walls — a notion that depends on the duration of the gas withdrawal.

When a constant temperature, T_0 , is applied at the flat surface of a semi-infinite space, the heat flux is $Q = \Sigma K_R (T_\infty - T_0) / \sqrt{\pi k_R t}$. Note that when, instead of a flat surface, an idealized spherical cavern with radius a is considered, the flux is $Q = \Sigma K_R (T_\infty - T_0) / (1/a + 1/\sqrt{\pi k_R t})$ [In software programs used by the cavern industry, a spherical cavern whose volume equals the actual cavern volume often is considered; its surface (together with the heat resistance coefficient \bar{H} , see Section 5.1) is selected to fit the actual measurements. Neither this procedure nor the procedure suggested in this paper can account both for short and long periods of time. The procedure used in this paper allows using the results of sonar surveys, which provide a good picture of cavern shape]. When gas temperature is not constant, $T = T(t)$, the flux can be expressed through a convolution:

$$Q(t) = K_R \Sigma \int_0^t \frac{\dot{T}(\tau)}{\sqrt{\pi k_R (t - \tau)}} d\tau \quad (17)$$

After some simplification, the energy equation can write:

$$\frac{\gamma}{\gamma - 1} \frac{P(t)}{T(t)} \dot{T}(t) - \dot{P}(t) = -\frac{\Sigma}{V} \frac{K_R}{\sqrt{\pi k_R}} \int_0^t \frac{\dot{T}(\tau)}{\sqrt{t - \tau}} d\tau + \frac{\langle \dot{m} \rangle}{V} C_p (T_{inj} - T) \quad (18)$$

When gas is withdrawn ($\dot{m} < 0$), temperature evolution results from gas expansion ($\dot{P} < 0$), whose effects are mitigated by the heat flux from the rock mass (right-hand side). Beside initial conditions, P_0 and T_0 , only two physical constants appear in this equation, $\gamma = C_p / C_v$ and $\Sigma K_R / V \sqrt{\pi k_R}$. The adiabaticity index, $\gamma = C_p / C_v$, equals $\gamma = 1.4$ (air, hydrogen) or $\gamma = 1.3$ (methane). A comparison of temperature evolution when air, methane or hydrogen is stored can be found in ³⁸Nieland, 2008, ⁶Schlichtenmayer and Bannach, 2015 and ³⁹Louvet et al., 2017. The influence of rock thermal properties occurs through the parameter $K_R / \sqrt{\pi k_R}$, whose value typically is $1800 \text{ W} \times \text{s}^{1/2} / \text{m}^2$ (see Table 2).

The ratio V/Σ depends on cavern shape and is proportional to cavern size. This is the parameter susceptible of the largest variations. Following a gas withdrawal, the temperature drop is larger in a bigger cavern. ⁴¹Staudtmeister et al., 2010, suggest $V/\Sigma = 8\text{-}12 \text{ m}$; larger values are found in a large cavern when the height/diameter ratio is close to 1. In an idealized spherical cavern with radius a , for instance, $V = 525,000 \text{ m}^3$, $a = 50 \text{ m}$ and

$V/\Sigma = a/3 = 17$ m. ⁴⁴Rokahr et al., 2011, mention that “the ratio V/Σ is usually less than 10 m for caverns in salt domes because they are often shaped long and thin” (p. 194).

5.9. Influence of the shape factor

An interesting example of the influence of the shape factor (Σ/V) was described by ¹⁵Quast (1983). The Huntorf CAES storage, mentioned in Section 3.2, consists of two caverns, NK1 and NK2, whose volumes accessible to air are $V_1 = 140\,000$ m³ and $V_2 = 170\,000$ m³, respectively (Fig. 14). Their pressure varies between 4.3 MPa and 7 MPa; maximum pressure rate during air withdrawal is $\dot{P} = 1.5$ MPa/hr. During operation, the two caverns are connected at ground level, which means that their pressures are almost perfectly equal. One can expect that the inventory is larger in the bigger cavern (NK2). ¹⁵Quast (1983) mentions that this was not observed when operating the caverns (injected masses were similar in the two caverns) as gas temperature increases more slowly in the NK1 cavern whose effective surface is larger.

5.10. The case of the Melville cavern

Instead of P, T in Eq. (18), m and T can be selected as the main variables, $m = PV_0/RT$:

$$mC_v\dot{T} - \dot{m}rT = -\frac{\Sigma K_R}{\sqrt{\pi k_R}} \int_0^t \frac{\dot{T}(\tau)}{\sqrt{t-\tau}} d\tau + \langle \dot{m} \rangle C_p(T_{inj} - T) \quad (19)$$

This integro-differential equation was tested against a withdrawal test ($\dot{m} < 0$) described by ⁴⁶Crossley, 1996. The test was performed in the Regina #5 natural-gas storage cavern at Melville (Canada). The measured flow rate, cavern pressure and temperature are drawn on Fig. 15. The withdrawal period was 5-days long. For computations (⁴⁷Djizanne et al., 2014), the following values were selected: $K_R = 6$ W/m/K, $\gamma = 1.305$, and $V = 46,000$ m³. Cavern shape was unknown, and the surface of the cavern walls was selected to be twice the surface of a sphere whose volume equals the actual cavern volume, $V/\Sigma = 3.6$ m. Note that slightly before the end of the withdrawal phase, gas begins to warm, as the withdrawal flowrate and the heat capacity of the gas mass are small, whereas the heat flux from the

rock mass is high. A similar observation was made by ⁴⁵Crotogino et al. (2001): “An interesting effect occurs when the air is expanded to atmospheric pressure [in the Huntorf caverns, see Sections 3.2 and 5.9]: the pressure drop initially results in the expected cooling of the air, but after reaching a minimum, the temperature rises again” (p. 355).

5.11. Gas warming in a shut-in cavern

In this section, we discuss a case similar to that discussed in Section 5.4. At $t = 0$, the cavern temperature, which was $T = T_\infty$, is decreased abruptly to $T = T_\infty - T_0$, and the cavern is shut in ($\dot{m} = 0$):

$$m_0 C_v \dot{T}(t) = -\frac{\Sigma K_R}{\sqrt{\pi k_R}} \left[\frac{T_0 - T_\infty}{\sqrt{t}} + \int_0^t \frac{\dot{T}(\tau)}{\sqrt{(t-\tau)}} d\tau \right] \quad (20)$$

As in Section 5.3, two characteristic times can be defined: a^2 / k_R and $a^2 / k_R \chi$, $\chi = \rho_R C_P^R / \rho_g C_P^g$, which are relative to the rock mass and the gas body, respectively. The latter characteristic time is much shorter than in the case of a liquid, and thermal equilibrium with the rock mass at the cavern wall is reached rapidly. However, in the first years following leaching completion, temperature at the cavern wall is much colder than the geothermal temperature at the cavern depth. The former characteristic time suggests that equilibrium temperature (between gas and rock wall) slowly increases with time (⁸Fosse and Røvang, 1998).

5.12. Thermodynamic evolution during a fast, small temperature change

When withdrawal is considered, a simplified solution of Eq. (18) can be reached as follows. It is assumed that temperature decreases by a small amount, $\delta T < 0$ (when compared to the initial temperature T_0), during a δt -long period and that temperature rate is constant; $\delta T / T$ is not very different from $\delta T / T_0$. (When $\delta T = 10$ °C, the relative error is 3%.) Then, Eq. (18) can be solved as follows:

$$\frac{\delta P}{\delta T} = \left(\frac{4 \Sigma K}{3 V \sqrt{\pi k}} (\delta t)^{1/2} + \frac{\gamma P_0}{\gamma - 1 T_0} \right) \quad (21)$$

Equation (21) proves that the temperature drop (for a given pressure drop) is larger when the initial pressure, P_0 is smaller when the pressure drop is faster and ratio Σ/V is smaller. This solution holds when δT and δP are not too large. For instance, $\gamma = 1.4$ (air), $T_0 = 310$ K, $K_R / \sqrt{\pi k_R} = 1800$ W/s^{1/2}m² lead to:

$$\frac{\delta P [\text{MPa}]}{\delta T [^\circ\text{C}]} = 0.072 \frac{\Sigma}{V} \delta t^{1/2} [\text{hr}^{1/2}] + 0.011 P_0 [\text{MPa}] \quad (22)$$

For instance, $P_0 = 7$ MPa, $\Sigma/V = 0.1$ m⁻¹, and $\delta T/\delta P = 9.5$ °C/MPa. Such a temperature drop can have significant structural consequences: tangential tensile stresses are generated at the cavern wall, and fracturing may occur. (This topic has been studied by many authors.) Orders of magnitude can be reached as follows. Consider the case of an idealized cylindrical cavern with radius a , cavern pressure P_c , and geostatic pressure, P_∞ . Cavern pressure is lowered from P_c to $P_c - \delta P$. The worst case (largest temperature drop) is reached when gas evolution is isentropic, $\delta P / \delta T = \gamma P_0 / (\gamma - 1) T_0$. Salt behavior is elasto-thermo-viscoplastic; a uniaxial formulation is: $\dot{\epsilon} = \dot{\sigma} / E - \alpha_R \dot{T} + A \sigma^n$. It is assumed that before the pressure drop, steady-state viscoplastic behavior is reached, and the vertical stress is $\sigma_{zz}(a) = -(P_\infty - P) / n$ (Wang et al., 2015). However, the effect of viscoplastic deformation is to lower deviatoric stresses and, considering the thermo-elastic deviatoric stresses is on the safe side. The vertical tangential stress immediately after a gas drop can written

$$\sigma_{zz}(a) + P = -\frac{P_\infty - P}{n} + \delta P - \frac{\bar{E}\bar{\alpha}}{1-\bar{\nu}} \delta T = -\frac{P_\infty - P}{n} + \delta P \left(1 - \frac{\bar{E}\bar{\alpha}}{1-\bar{\nu}} \frac{\gamma-1}{\gamma} \frac{T_0}{P_0}\right)$$

Tensile effective stresses, $\sigma_{zz}(a) + P > 0$, can appear ($\bar{E} = 25$ GPa, $\bar{\alpha} = 4 \times 10^{-4}$ / °C, $\bar{\nu} = 0.3$, $\gamma = 1.4$, $T_0 = 300$ K and $P_0 = 7$ MPa are typical of an air storage, and $\bar{E}\bar{\alpha}T_0(\gamma-1) / \gamma P_0(1-\bar{\nu})$ is much larger than 1.) and can lead to fracturing (Sicsic and Bérest, 2014).

6. Gas flow in the wellbore; hydrate formation

Especially during rapid injection or withdrawal, gas temperature and pressure are not identical at ground level ($z = 0$) and at the cavern entrance ($z = H$) because of head losses and gravity forces. In the case of gas production, this problem was discussed through numerical

computations (for instance, by ⁵⁰Kabir et al., 1996). When gas contains water vapor during withdrawal, hydrates can form in the wellbore when a certain relation between gas pressure, temperature and moisture content is met, $c_v > c_{hydrate}(P_w, T_w)$. Hydrates are harmful for surface equipment and pipes. Most often, an inhibitor (methanol) is injected at shallow depth in the wellbore to prevent hydrate formation. This operation is costly, and various solutions have been proposed in the literature to hinder hydrate formation (⁵¹Réveillère et al., 2016). In fact, thermodynamic evolution can be computed when several simplifications are accepted. The average gas speed is in the $u = 1-15$ m/s range; the wellbore diameter is $D = 0.2-0.3$ m; and the gas viscosity is $\bar{\nu}_g = 10^{-7}$ m²/s (when gas density is $\rho_g = 200$ kg/m³). The Reynolds number, $R = Du / \bar{\nu}$, is larger than 10^4 , gas flow is turbulent, and the gas speed is uniform, $u = u(z,t)$, in any horizontal section (except for a thin boundary layer at the casing wall). When the gas speed is $u = 10$ m/s, 100 s are needed for gas to cross through a wellbore that is $H = 1000$ -m long. During such a period of time, cavern pressure and temperature are not given time enough to change significantly: steady-state is reached, and mass conservation can be written $\dot{m} = S\rho(z)u(z)$, where \dot{m} is a slowly varying function of time, and S is the wellbore cross-sectional area. (In fact, wellhead temperature increases first during a withdrawal, as the cavern gas is warmer than that at ground level (⁵²Grosswig et al., 2001), a transient phenomenon.) Along the wellbore, rock temperature does not equal gas temperature, and heat transfer takes place. However, during a one-day period during which 10% of the gas inventory is withdrawn, rock temperature experiences significant changes in a cylinder whose radius is $a = \sqrt{k_R t} = 0.5$ m and whose height is $H = 1000$ m. Its thermal capacity is $\pi a^2 H \rho_R C_P^R$, where $\rho_R C_P^R = 2 \times 10^6$ J/m³/°C. The cavern volume is $V_c = 500\,000$ m³, of which $\delta V_c = 10\%$ is withdrawn; the heat capacity of the withdrawn natural gas is $\delta V_c \rho_g C_P^g$, where $\rho_g C_P^g = 4 \times 10^5$ J/m³/°C (Gas pressure is $P_c = 20$ MPa.), and $\pi a^2 H \rho_R C_P^R = 1.5 \cdot 10^9$ J/m³/°C is much smaller than $\delta V_c \rho_g C_P^g = 2.25 \cdot 10^{11}$ J/m³/°C. The change in rock temperature (less than a few dozen °C) is much larger than the gas temperature change, and gas flow is considered as adiabatic. The set of equations for speed, density, pressure and entropy can be written

$$\dot{m} = S \rho_w(u) u \quad (23)$$

$$\frac{\gamma}{\gamma-1} \frac{P_w(u)}{\rho_w(u)} + \frac{u^2}{2} - gz = I \quad (24)$$

$$\frac{1}{\rho_w(u)} \frac{dP_w[u(z)]}{dz} + u \frac{du}{dz} - g = -\varepsilon F u^2 \quad (25)$$

$$u \frac{s_w[u(z)]}{dz} \geq 0 \quad (26)$$

Equation (23) reflects mass conservation. Equation (24) is the enthalpy equation, where g is gravity acceleration, and I is a constant along a flow path that can be computed at $z = 0$, $u = u_0$ or at $z = H$, $u = u_c$. Equation (25) is the momentum equation, where $F u^2 > 0$ is the head loss per unit of length, F is a function of duct diameter, wall roughness, etc., and $\varepsilon = \pm 1 = \text{sgn}(u)$. Entropy inequality, Eq. (26), also can be written as follows: $dh(s, \rho) = T ds + dP / d\rho|_s d\rho$, $dP / d\rho|_s = c^2$, where c is the velocity of sound ($c = \sqrt{\gamma r T}$; for methane, $c = 450$ m/s is typical), and $dh = u du - g dz$ or $T ds = (c^2 - u^2) du / u + g dz$. The gas speed must be smaller than the velocity of sound, a condition that is satisfied when u equals a few dozen m/s. In the case of a blowout, u can equal c ; this case is discussed in ⁵³Djakeun-Djizanne et al., 2013. Only standard gas speed during operation are considered here. From Eqs. (23, 24, 25), P and ρ can be eliminated, leading to

$$-\frac{\gamma+1}{2(\gamma-1)} + \left[\frac{gz + I}{u^2} + \frac{1}{\gamma-1} \left(\frac{g}{u} - \gamma \varepsilon F u \right) \frac{dz}{du} \right] = 0 \quad (27)$$

When gravity is disregarded ($g = 0$), this equation can be solved easily (⁵³Djakeun-Djizanne et al., 2013):

$$\left[-\frac{\gamma+1}{4(\gamma-1)} \text{Log} v^2 - \frac{I}{v^2} \right]_{v=u_0}^{v=u} - \frac{\gamma}{\gamma-1} \varepsilon F z = 0 \quad (28)$$

When gravity is taken into account, the solution can be written

$$\left[(gz(u) + I) \left(\frac{u}{\sqrt{|1 - \gamma \varepsilon F u^2 / g|}} \right)^{\gamma-1} \right]_{z=0}^{u(z)} = \frac{\gamma+1}{2} \int_{u_0}^{u(z)} \left(\frac{v}{\sqrt{|1 - \gamma \varepsilon F v^2 / g|}} \right)^{\gamma-1} \frac{v dv}{1 - \gamma \varepsilon F v^2 / g} \quad (29)$$

from which $u(z)$, $\rho_w(u)$, $T_w(u)$ and $P_w(u)$ can be inferred. The withdrawal case ($\varepsilon = -1$, $\dot{m} < 0$, $u < 0$) is simple when \dot{m} is fixed, and boundary conditions $\rho(H) = \rho_c$, $P(H) = P_c$,

$T(H) = T_c$ and $I = I_c$ are known at cavern depth. Computations are straightforward. The moisture content such that hydrate can form is a function of pressure and temperature. When moisture content at the cavern top (from which gas is withdrawn) is known, hydrate formation at shallow depth in the wellbore can be discussed. Computations allow discussion of hydrate formation, which is a function of water vapor concentration in the cavern, and gas pressure and temperature in the well (Klafki et al., 2003). The case of gas injection ($\varepsilon = 1, \dot{m} > 0, u > 0$) is slightly more complicated: $P(H) = P_c$ is known, but neither T nor ρ is continuous through $z = H$; a second boundary equation (at ground level, $z = 0$) must be used. In addition, du/dz can vanish when $Fu^{*2} = \gamma g$, which means that $u = u(z)$ is not a monotonous function of z .

Conclusion

Various aspects of the thermodynamic behavior of gas- and liquid-filled caverns were discussed. In the case of liquid-filled caverns, onset of natural convection is certain. Rapid pressure changes (1 day) lead to small adiabatic temperature changes. Conversely, thermal equilibrium between the cavern liquid and the rock mass is reached after a long period of time, raising the problem of long-term behavior after cavern decommissioning. The heat capacity of gas-filled caverns is much smaller, and most features of thermodynamic behavior are different. Natural convection is not always present, and the reasons for this are not perfectly clear, although water vapor distribution in the cavern might be an important factor. Gas evolution is not adiabatic, even during short injections/withdrawals. During withdrawal, gas is cooled in the cavern and in the wellbore, an evolution that can be captured through simple equations. Gas cooling leads to possible tensile fracture creation at the cavern wall and hydrate formation at wellbore top. Most of these features can be captured by simple models. Advances in the understanding of cavern behavior will rely mostly on field data, which are relatively scarce.

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Table headings

Table 1. Gas thermal constants, from ⁵Gas Encyclopedia, Air Liquide. Constants are measured at $\theta = 25^{\circ}\text{C}$ and $P_N = 0.10325 \text{ MPa}$.

Table 2. Values of $K_R / \sqrt{\pi k_R}$

Figure captions

Figure 1. Temperature profile along the vertical axis of the EZ53 brine cavern from the Etrez site in France (¹Brouard, 1998, courtesy of Storengy).

Figure 2. As-computed convection cells in the EZ53 cavern, assuming axisymmetry. Left: iso-values of the streamline function; right: velocity magnitude (²Karimi-Jafari, 2007).

Figure 3. Temperature profile along the vertical axis of symmetry in a brine cavern (³Banach and Klafki, 2009).

Figure 4. Temperature profile in the S107 natural gas cavern, Stassfurt, Germany (left); cavern shape selected for numerical computations (right) (after ⁴Klafki et al., 2003).

Figure 5. Temperature profile in a cavern at Etzel, Germany; roof depth is slightly more than 1200 m, and the gas interface is slightly above 1500 m (¹³Thaule and Gentzsch, 1994).

Figure 6. The cavern profile (1) is drawn on the left side of each picture. In the left illustration, temperature profiles before (5) and after (2, 3, 4) three consecutive withdrawals are represented; the right illustration represents how temperature reaches thermal equilibrium after the withdrawals are completed (2, 3) (⁴Klafki et al., 2003).

Figure 7. Gas flowrates (above); for air, $1 \text{ Nm}^3 = 1.29 \text{ kg}$. Temperature profiles along a vertical axis in the NK1 cavern of the Huntorf CAES after injections and withdrawals (below) (¹⁵Quast, 1983).

Figure 8. Enterprise Cavern 1 (left) — At a 3751-ft (1144-m) depth, gas pressure was decreased from 2728 psi (18.8 MPa) to 1994 psi (13.7 MPa) between April 2009 and May 2009 (¹⁶Skaug et al., 2010); (right) a tentative explanation.

Figure 9. Temperature profile in an anonymous gas cavern: the temperature gradient in the upper part of the cavern is $G = 6 \times 10^{-3} \text{ }^\circ\text{C/m}$; it is negative above the sump. (After ¹⁷Berger et al., 2002).

Figure 10. Temperature evolution in an 8000-m³ cavern after solution mining is completed (Courtesy of Storengy).

Figure 11. Pressure evolution in three caverns from the same site whose sizes are different (²⁹Van Sambeek et al., 2005; courtesy of Storengy).

Figure 12. Pressure and Temperature evolutions in an oil cavern during injections/withdrawals (³⁴Gatelier et al., 2008).

Figure 13. Cavern temperature as a function of cavern pressure during a cycle in a gas storage cavern in the Gulf of Mexico. The blue dotted line represents the computed adiabatic behavior $P^{ad} / P_0 = (T^{ad} / T_0)^{1-\gamma}$ (after ³⁵Heath and Benefield, 2010).

Figure 14. Vertical cross sections of Huntorf NK1 (left) and NK2 (right) caverns. From ⁴⁵Crotogino et al. 2001.

Figure 15. Melville Cavern: Gas withdrawal rate, pressure and temperature evolutions as observed (after ⁴⁶Crossley, 1996) and as computed: 1 MSCFH = 28,317 Nm³/hr, 1 MPa = 145 psig and 20° C = 68° F.

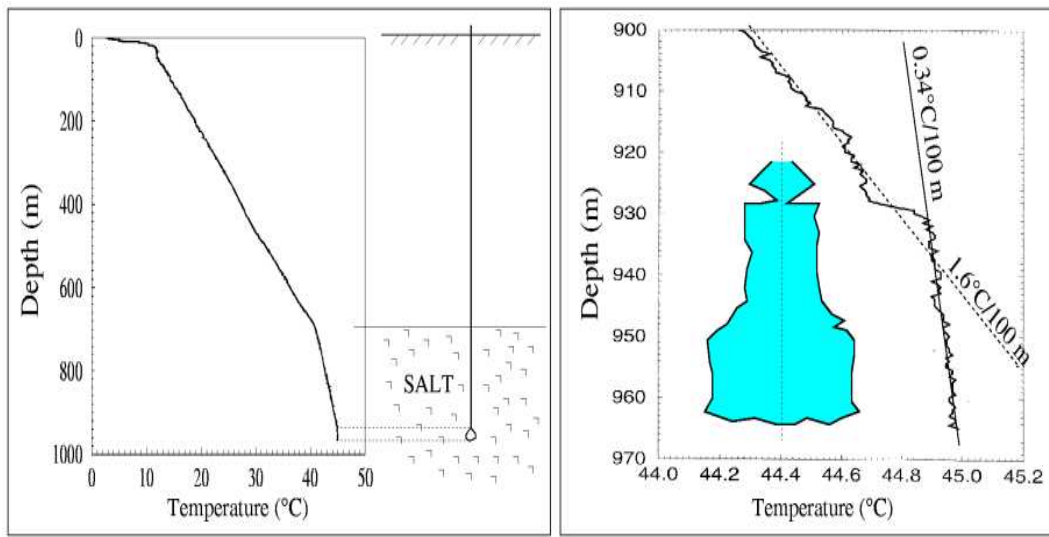


Figure 1. Temperature profile along the vertical axis of the EZ53 brine cavern from the Etrez site in France (¹Brouard, 1998, courtesy of Storengy).

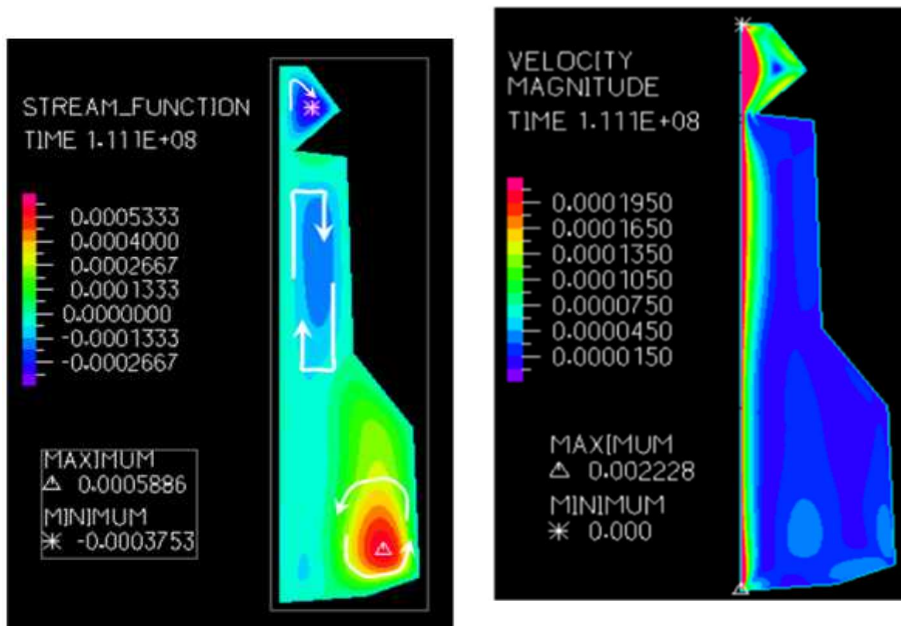


Figure 2. As-computed convection cells in the EZ53 cavern, assuming axisymmetry. Left: iso-values of the streamline function; right: velocity magnitude (²Karimi-Jafari, 2007).

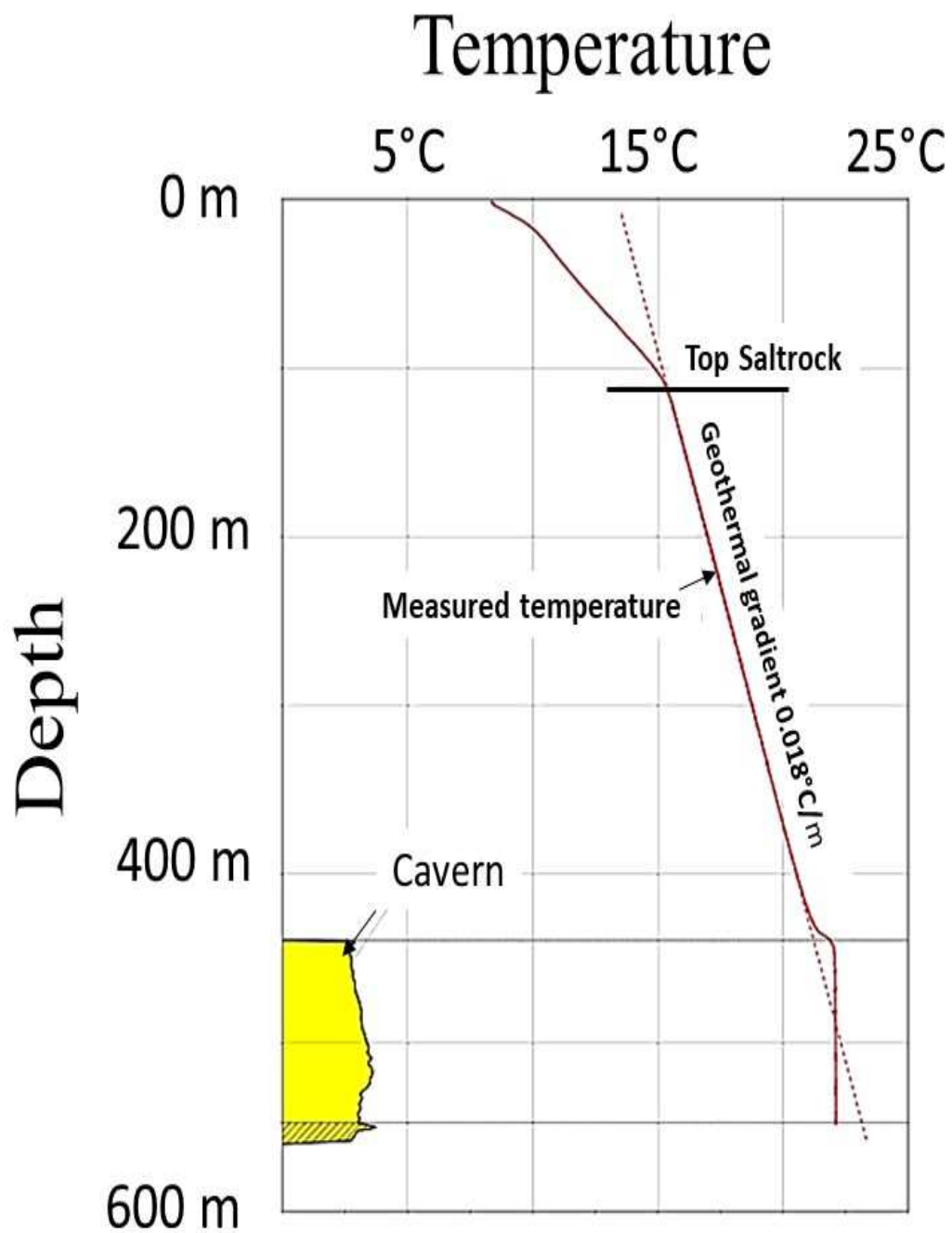


Figure 3. Temperature profile along the vertical axis of symmetry in a brine cavern (After ³Banach and Klafki, 2009).

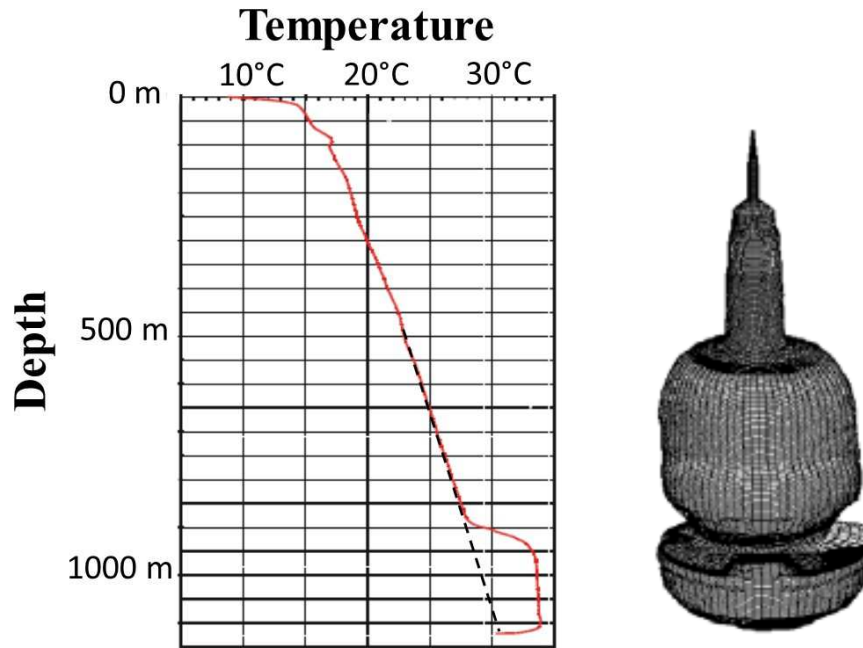


Fig. 4. Temperature profile in the S107 natural gas cavern, Stassfurt, Germany (left); cavern shape selected for numerical computations (right) (after ⁴Klafki et al., 2003).

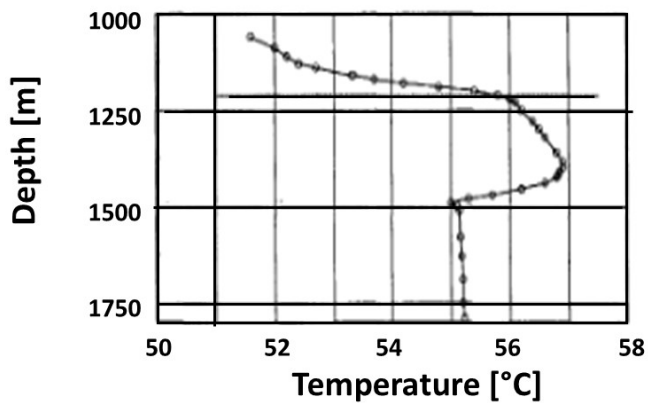


Figure 5. Temperature profile in a cavern at Etzel, Germany; roof depth is slightly more than 1200 m, and the gas interface is slightly above 1500 m (After ¹³Thaule and Gentzsch, 1994).

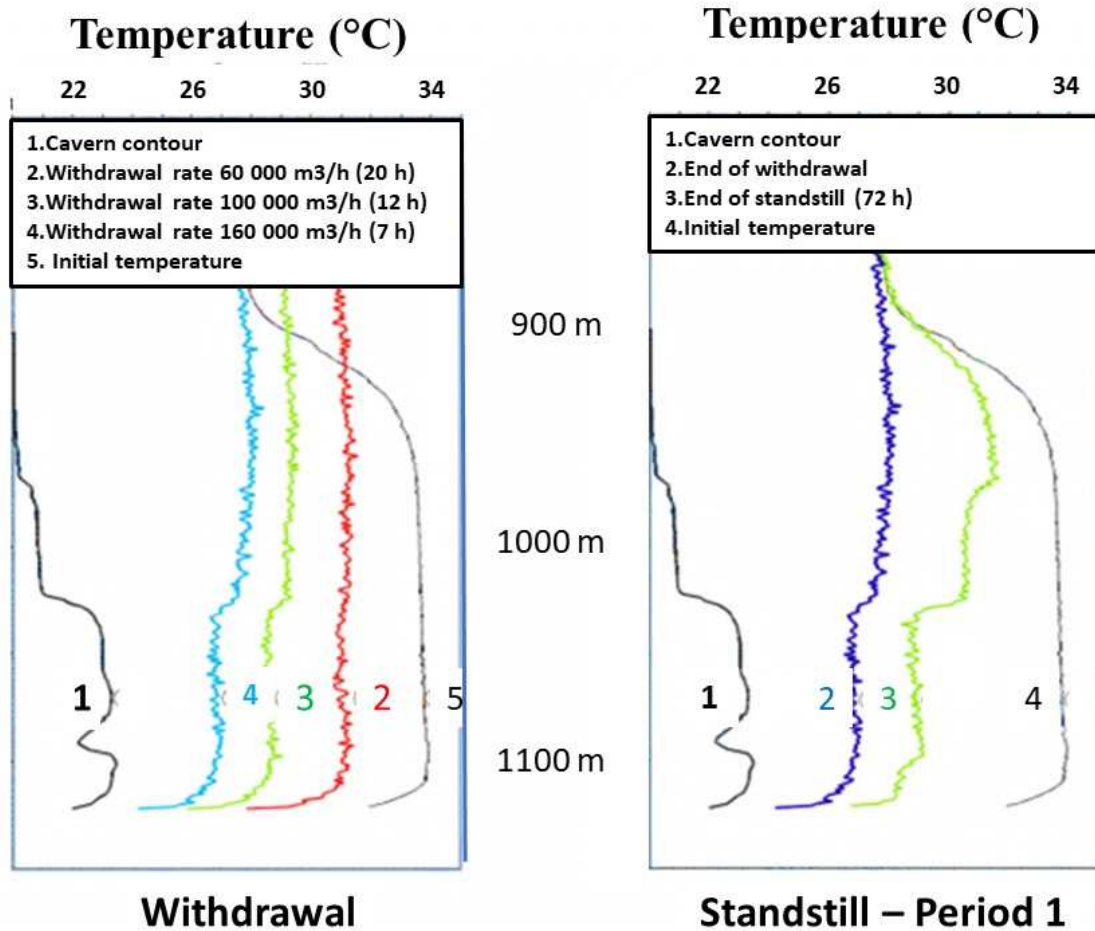


Figure 6. The cavern profile (1) is drawn on the left side of each picture. In the left illustration, temperature profiles before (5) and after (2, 3, 4) three consecutive withdrawals are represented; the right illustration represents how temperature reaches thermal equilibrium after the withdrawals are completed (2, 3) (After ⁴Klafki et al., 2003).

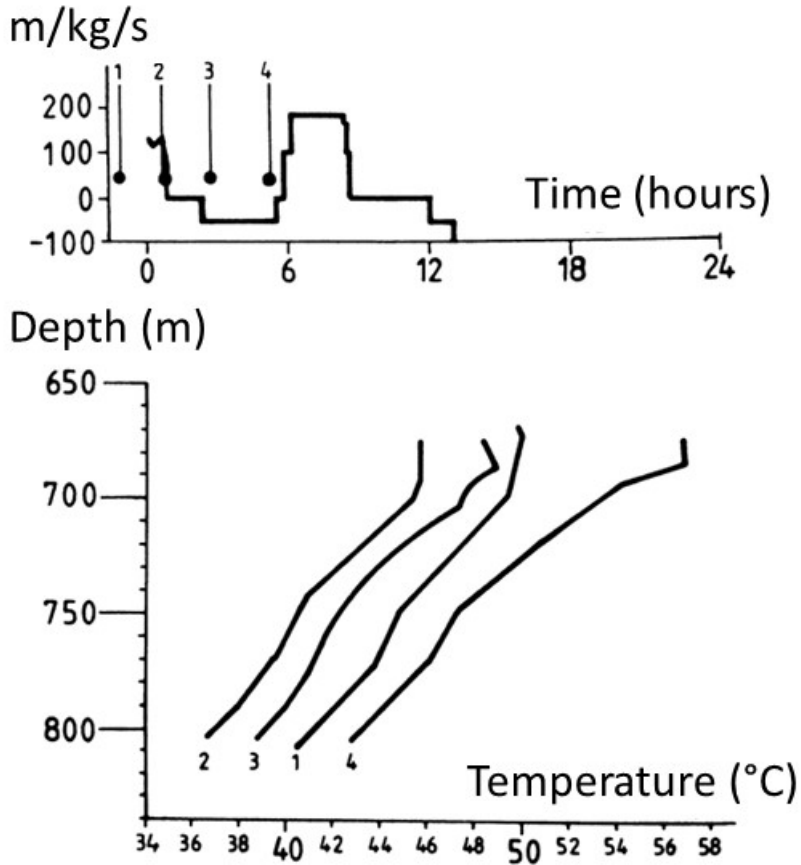


Figure 7. Gas flowrates (above); for air, $1 \text{ Nm}^3 = 1.29 \text{ kg}$. Temperature profiles along a vertical axis in the NK1 cavern of the Huntorf CAES after injections and withdrawals (below) (After ¹⁵Quast, 1983).

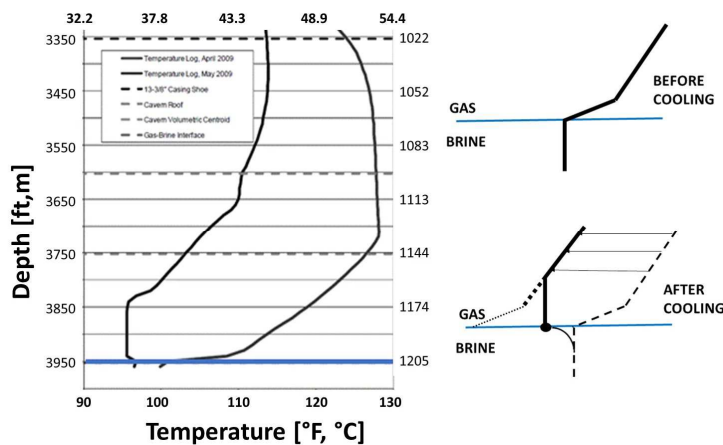


Figure 8. Enterprise Cavern 1 (left) — At a 3751-ft (1144-m) depth, gas pressure was decreased from 2728 psi (18.8 MPa) to 1994 psi (13.7 MPa) between April 2009 and May 2009 (After ¹⁶Skaug et al., 2010); (right) a tentative explanation.

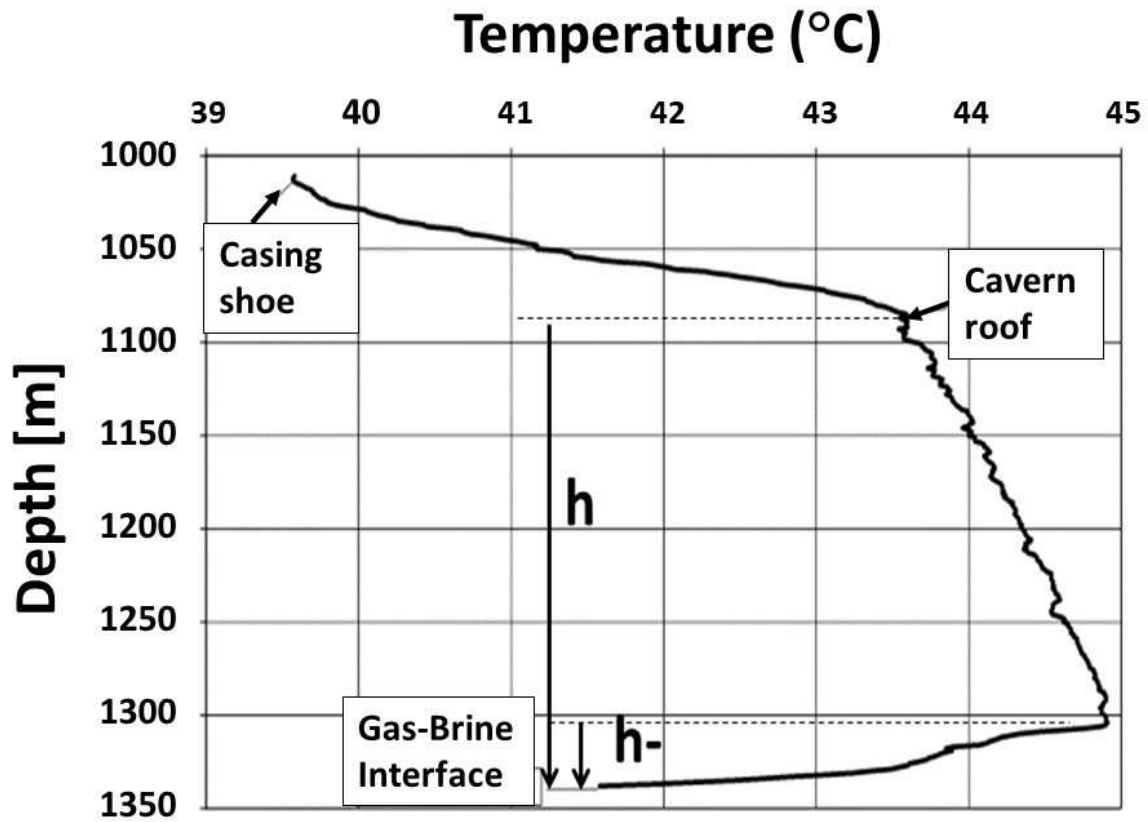


Figure 9. Temperature profile in an anonymous gas cavern: the temperature gradient in the upper part of the cavern is $G = 6 \times 10^{-3} \text{°C/m}$; it is negative above the sump. (After

¹⁷Berger et al., 2002).

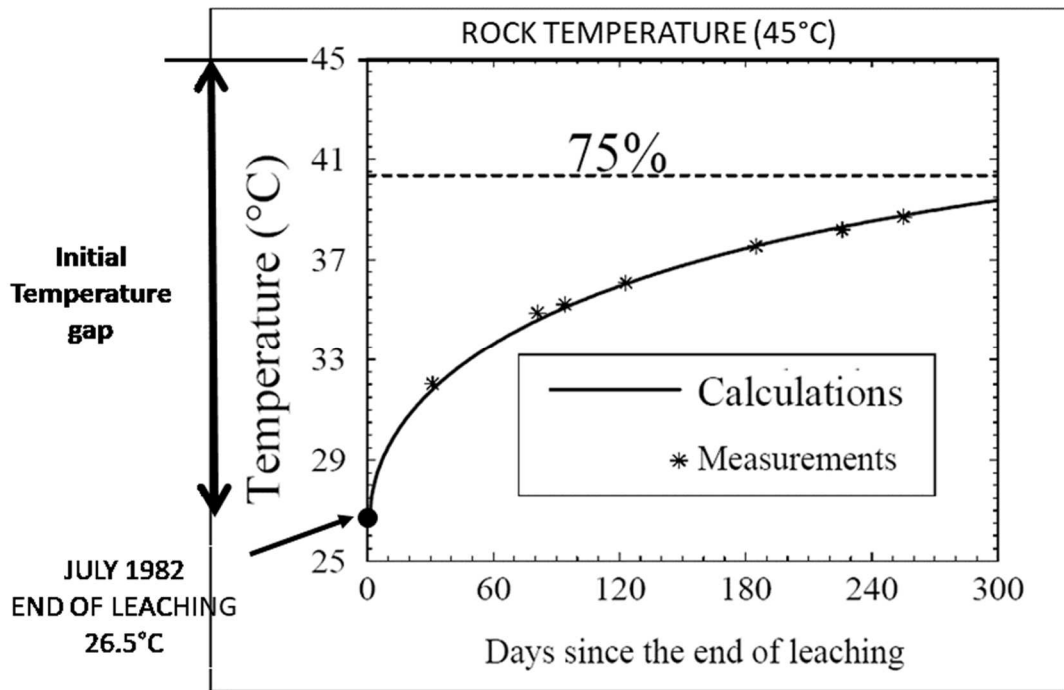


Figure 10. Temperature evolution in an 8000-m³ cavern after solution mining is completed (Courtesy of Storengy).

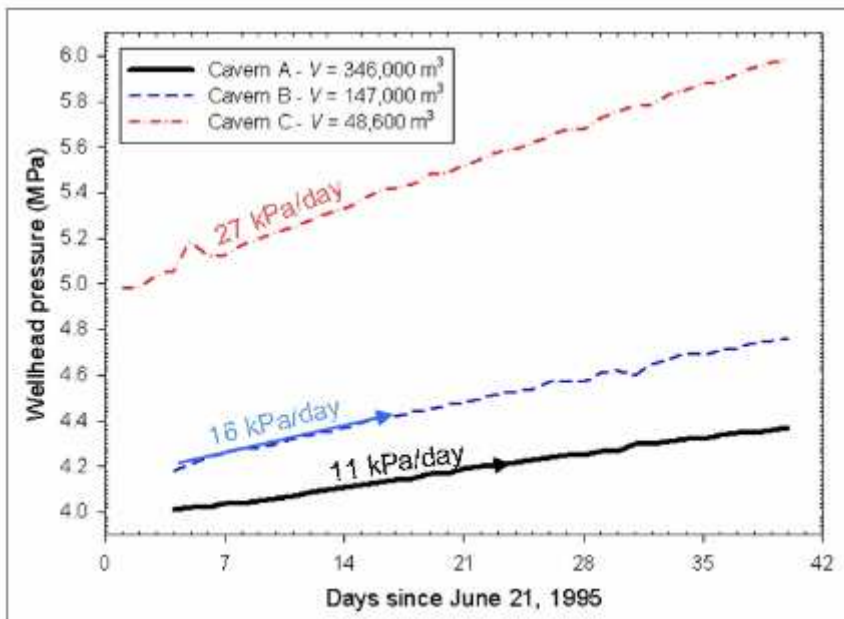
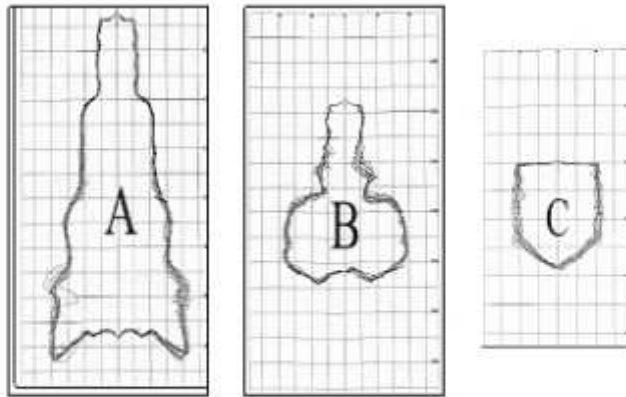


Figure 11. Pressure evolution in three caverns from the same site whose sizes are different (³⁰Van Sambeek et al., 2005; courtesy of Storengy).

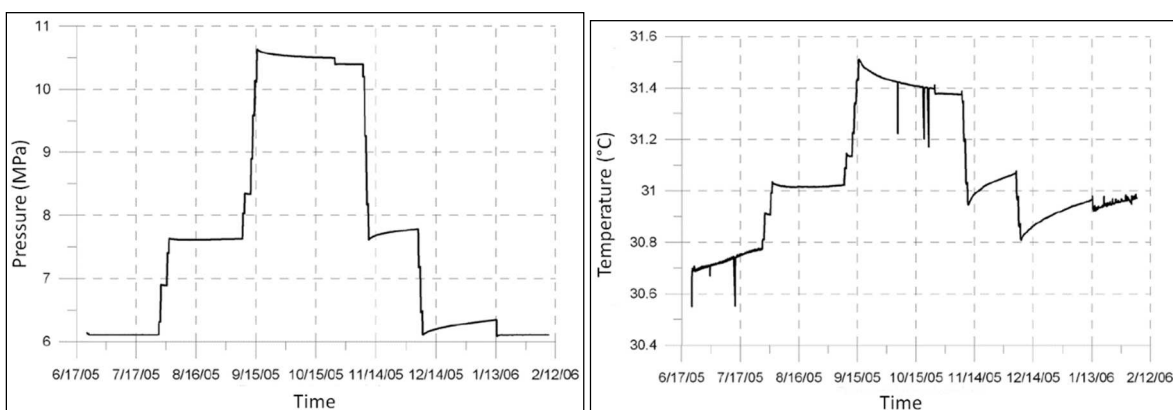


Figure 12. Pressure and Temperature evolutions in an oil cavern during injections/withdrawals (³⁴Gatelier et al., 2008).

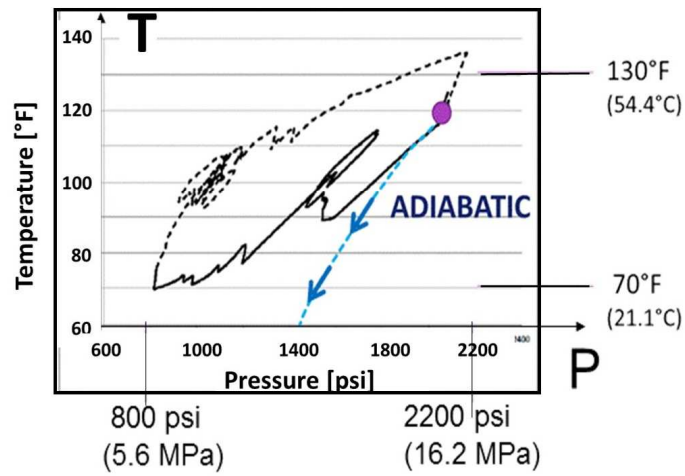


Figure 13. Cavern temperature as a function of cavern pressure during a cycle in a gas storage cavern in the Gulf of Mexico. The blue dotted line represents the computed adiabatic behavior $P^{ad} / P_0 = (T^{ad} / T_0)^{1-\gamma}$ (after³⁵Heath and Benefield, 2010).

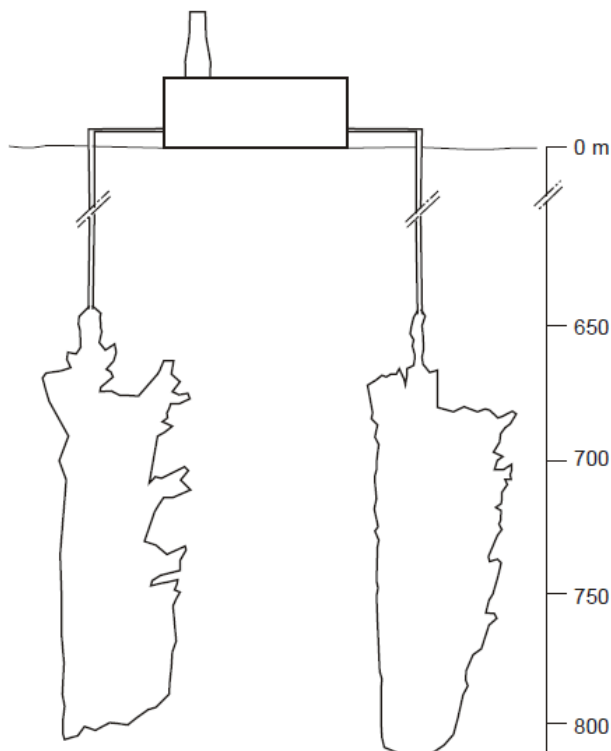


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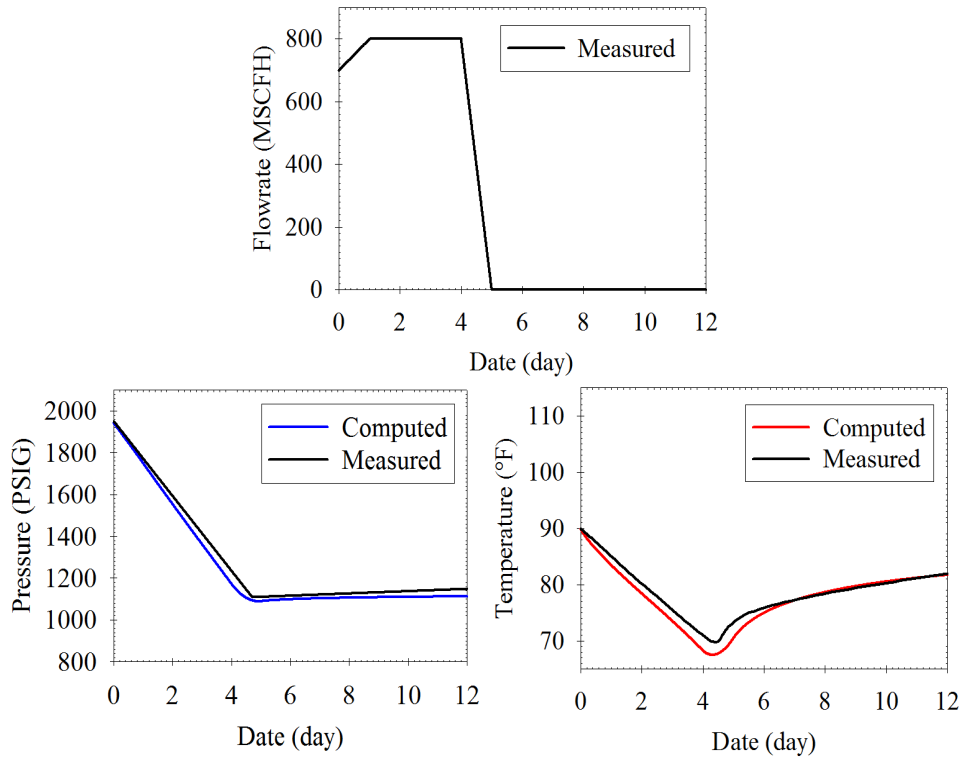


Figure 15. Melville Cavern: Gas withdrawal rate, pressure and temperature evolutions as observed (after ⁴⁶Crossley, 1996) and as computed: 1 MSCFH = 28,317 Nm³/hr, 1 MPa = 145 psig and 20° C = 68° F.

Table 2. Values of $K_R / \sqrt{\pi k_R}$

	K_{salt} W/m-°C	ρ_{salt} kg/m ³	C_{salt} J/kg-°C	k_{salt} m ² /s	$K_{salt} / \sqrt{\pi k_{salt}}$ SI
1	5.2	2174	800	2.99×10^{-6}	1700
2	5.5	2180	870	2.9×10^{-6}	1820
3	5	2200	860	2.64×10^{-6}	1740
4	6.1	2200	921	3.0×10^{-6}	1980

(1 = ⁴⁰Pellizzaro et al., 2011; 2 = ⁴¹Staudtmeister et al., 2011
3 = ⁴²Blanco Martin et al., 2015; 4 = ⁴³Brouard Consulting, 2019.

Table 1. Gas thermal constants, from ⁵Gas Encyclopedia, Air Liquide. Constants are measured at $\theta = 25^\circ\text{C}$ and $P_N = 0.10325\text{ MPa}$.

Gas	C_p (J/kg/K)	C_v (J/kg/K)	γ (-)	μ (Pa.s)	K_F (W/m/K)	$Pr = \mu C_p / K$	$G_{ad} = g / C_p$
Air (25 °C)	1006	718	1.402	1.84×10^{-5}	26×10^{-3}	0.7	9.8×10^{-3}
CH ₄ (25 °C)	2232	1708	1.306	1.1×10^{-5}	34×10^{-3}	0.72	4.4×10^{-3}
H ₂	14306	10180	1.405	0.89×10^{-5}	185×10^{-3}	0.69	0.7×10^{-3}
O ₂	920	658	1.397	2.06×10^{-5}	26×10^{-3}	0.73	10.6×10^{-3}
CO ₂	851	657	1.29	1.49×10^{-5}	16×10^{-3}	0.79	11.5×10^{-3}