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Valentin Quesneau, Wenqian Shan, Nicolas Desbois, Stéphane Brandes, Yoann Rousselin, et al.. Cobalt Corroles with Bis-Ammonia or Mono-DMSO Axial Ligands. Electrochemical, Spectroscopic Characterizations and Ligand Binding Properties. European Journal of Inorganic Chemistry, 2018, 2018 (38), pp.4265-4277. 10.1002/ejic.201800897 . hal-03481809

HAL Id: hal-03481809

<https://hal.science/hal-03481809>

Submitted on 15 Dec 2021

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Cobalt Corroles with Bis-Ammonia or Mono-DMSO Axial Ligands. Electrochemical, Spectroscopic Characterizations and Ligand Binding Properties

Valentin Quesneau,^[a] Wenqian Shan,^[b] Nicolas Desbois,^[a] Stéphane Brandès,^[a] Yoann Rousselin,^[a] Meddy Vanotti,^[c] Virginie Blondeau-Patissier,^[c] Mario Naitana,^[a]

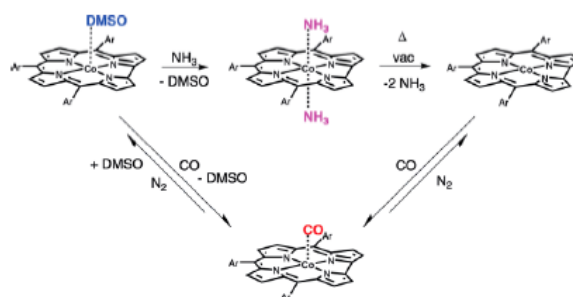
Paul Fleurat-Lessard,^[a] Eric Van Caemelbecke,^[b,d] Karl M. Kadish,^{*,[b]} and Claude P. Gros^{*,[a]}

Abstract: Four bis-ammonia ligated cobalt corroles and four mono-DMSO ligated cobalt corroles with different *meso*-aryl substituents on the macrocycle (A₂B- and A₃-corroles) were synthesized and investigated as to their electrochemical and spectroscopic properties under different solution conditions. The complexation energies of the investigated cobalt corroles were theoretically calculated to illustrate the propensity of the cobalt center for penta-coordination or hexa-coordination with various axial ligands (DMSO, CO, py and NH₃). The structure of one hexacoordinate bis-NH₃ cobalt corrole complex was also determined by X-ray diffraction.

Introduction

Metalloporphyrins are now an often-studied group of macrocycles because of their unique spectral and electrochemical properties and their potential applications as catalysts for a variety of reactions. Although corroles show many similarities with porphyrins, their distinctive structural, spectroscopic and photophysical properties as well as their specific coordination chemistry and particular chemical reactivity places these macrocycles in a distinctive position among members of the porphyrinoid family.^[1] Particularly, cobalt corroles exhibit a rich redox chemistry which depends on both the type of macrocycle and axial coordination at the cobalt metal center.^[1b] We have shown that solvents such as DMSO or pyridine as well as diatomic molecules such as CO will strongly coordinate to the cobalt center of the neutral corrole, leading not only to changes in electron density at the metal ion, but also in some cases to a change of metal and ligand oxidation state.^[2] Electron-withdrawing groups anchored at the periphery of the macrocyclic core will also change electron density at the cobalt central metal ion, thus directly affecting the magnitude and degree of axial coordination.^[3]

Recently, we reported the synthesis and electrochemical characterization of nitrophenyl-substituted cobalt corroles axially ligated with DMSO or pyridine molecules in their air-stable form.^[2,3] The current manuscript is an expansion of these studies and characterizes in greater detail the properties of cobalt corroles with a single DMSO axial ligand or two NH₃ axial ligands in their air-stable form. It is worthy to note that the hexacoordinate bis-ammonia cobalt corroles can easily be converted into the tetra-coordinated derivatives as shown in Scheme 1. Removal of the two coordinated ammonia molecules can be accomplished either under vacuum at 80 °C or by dissolving the complex in dichloromethane and bubbling N₂ gas through the solution. A schematic representation of the investigated compounds is given in Scheme 2.



Scheme 1. From pentacoordinate to hexacoordinate cobalt corroles; easy access to the tetra-coordinated species.

We have been interested in the selective binding and detection of trace amounts of CO at the sub-ppm level in the atmosphere by using Surface Acoustic Wave (SAW) sensors functionalized by cobalt corroles.^[4] The pathway in Scheme 1 allows for a smooth and easy access to the active species (the tetra-coor-

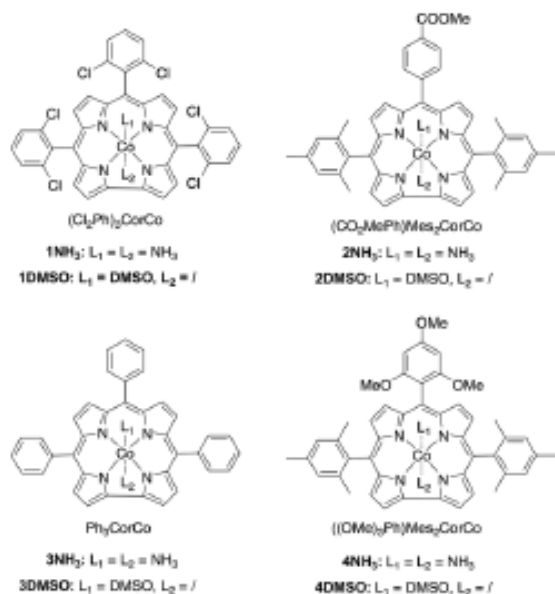
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Scheme 2. Structures of the investigated bis-ammonia and mono-DMSO meso-aryl cobalt corroles.

minated form of A_3 - and/or A_2B -cobalt corroles) which will bind CO. Very recently, concentrations of CO as low as a few hundred ppb were measured by SAW sensors emphasizing the interest of such sensors for the detection of carbon monoxide.^[4]

Results and Discussion

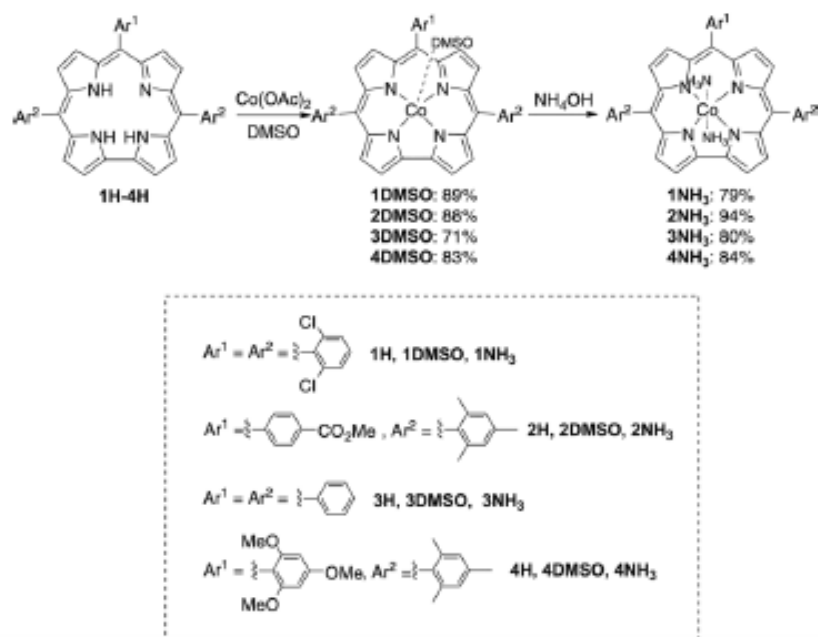
Synthesis of Cobalt Corrole Complexes

The eight cobalt corroles **1DMSO–4DMSO** and **1NH₃–4NH₃** were prepared according to the synthetic procedure shown in Scheme 3. The A_3 - and A_2B -free-base corroles **1H–4H** were synthesized according to published procedures described by Gryko and co-workers,^[3] and metallated with $\text{Co}(\text{OAc})_2$ in DMSO to give the cobalt complexes **1DMSO–4DMSO**, each with a single DMSO molecule as axial ligand. These mono-DMSO adducts were then treated with an aqueous ammonia solution to give the bis-ammonia derivatives **1NH₃–4NH₃** in 79–94 % yield.

Spectroscopic Characterization

The degree of axial coordination at the cobalt corrole center can be monitored by UV/Vis spectroscopy where large differences are seen between the four, five and six-coordinate forms of the compound. For example, as shown in Figure 1 for **1CO** and **1DMSO**, the two pentacoordinate cobalt corroles are characterized by a moderate absorption at 553 and 570 nm in the visible region of the spectrum. This spectral pattern contrasts with both the six coordinate **1NH₃**, that is characterized by an intense absorption band at 616 nm and a split Soret band at 438 and 452 nm and the tetra-coordinate derivative **1** which lacks axial ligands and displays a much blue-shifted Soret band at 390 nm as seen in Figure 1.

In order to examine the lability of the DMSO axial ligands in the currently examined corroles, the UV/Visible spectra of



Scheme 3. Synthesis of mono-DMSO and bis-NH₃ cobalt corroles.

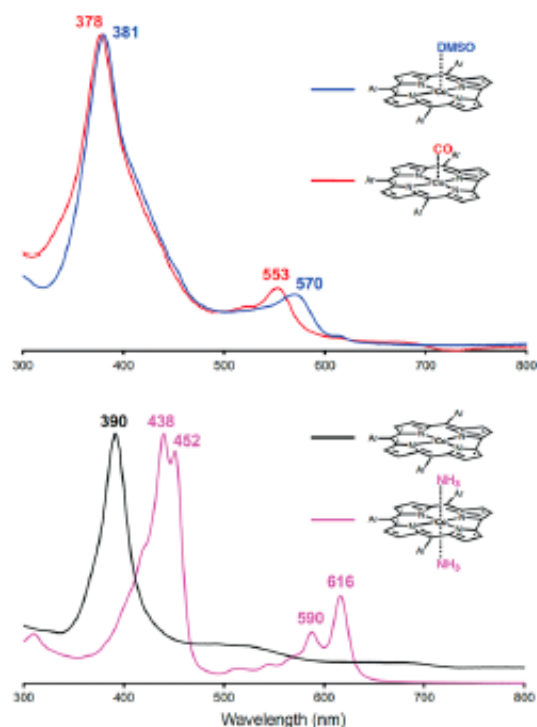


Figure 1. UV/Visible spectra (in CH_2Cl_2) of cobalt corrole **1** as a function of the axial coordination (DMSO, NH_3 , CO or nothing).

1DMSO–4DMSO were measured at a concentration of 10^{-3} M in both CH_2Cl_2 and DMSO containing 0.1 M TBAP (Figure 2). In CH_2Cl_2 the spectra are characterized by a Soret band at 379–387 nm and a single less intense band at 562–566 nm. There is also a shoulder on the Soret band at about 414–417 nm. These spectra are similar to previously reported UV/Vis spectra for a related series of mono-DMSO ligated cobalt nitrophenylcorroles under the same solution conditions.^[2]

Similar spectral patterns are seen for **2DMSO**, **3DMSO** and **4DMSO** in the two solvents but this is not the case for **1DMSO** as seen in Figure 2. The spectrum in CH_2Cl_2 (Figure 2a) is assigned to the mono-DMSO adduct while in DMSO a mixture of the five- and six-coordinate derivatives are proposed to exist, the hexa-coordinate DMSO adduct having bands at 417, 442, 453 and 618 nm as described on the following pages. It should be noted that **1DMSO** is the only cobalt corrole among the currently investigated compounds which possesses six electron-withdrawing substituents on the *meso*-phenyl rings, and this seems to facilitate the binding of a second DMSO molecule to the cobalt center of the neutral compound. Stronger pyridine binding constants were also earlier reported for cobalt corroles with electron-withdrawing substituents on the *meso*-phenyl rings.^[2]

UV/Visible spectra of the bis-ammonia complexes **1NH₃–4NH₃** were also measured at a concentration of 1×10^{-3} M un-

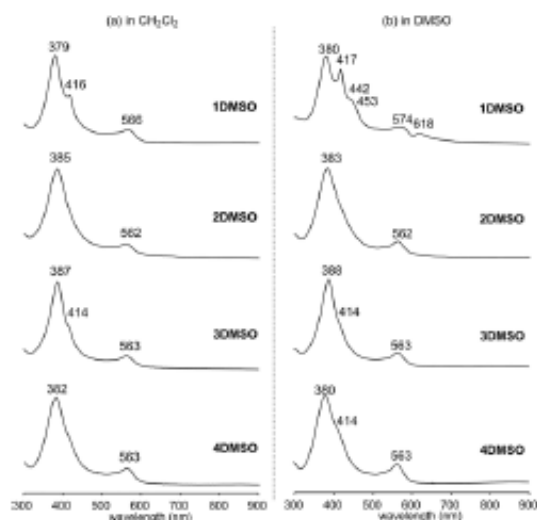


Figure 2. UV/Visible spectra of **1DMSO–4DMSO** at 10^{-3} M concentration in (a) CH_2Cl_2 and (b) DMSO containing 0.1 M TBAP.

der two solution conditions, one being CH_2Cl_2 containing 0.1 M TBAP and the other CH_2Cl_2 containing 0.1 M TBAP and 1 % (w/v) NH_4OH which produces NH_3 in solution. Examples of the corrole spectra in the two solvents are shown in Figure 3 and a summary of the spectroscopic data for **1NH₃–4NH₃** in CH_2Cl_2 containing 1 % NH_4OH is given in Table 1. As seen in Figure 3, b, the UV/Visible spectrum of each bis- NH_3 derivative is characterized by a split Soret band and two or three visible bands,

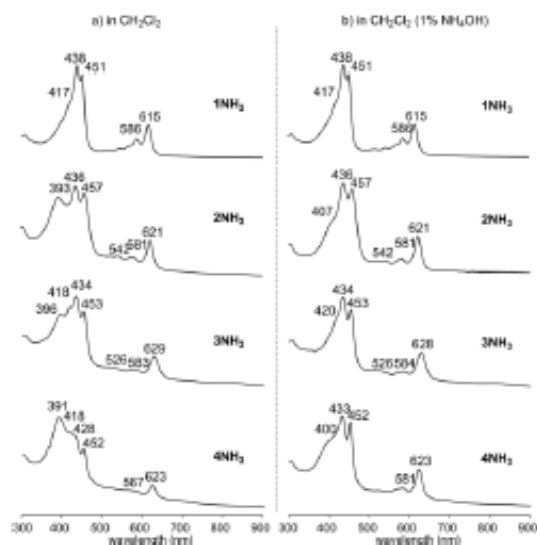


Figure 3. UV/Visible spectra of **1NH₃–4NH₃** at 10^{-3} M concentration in a) CH_2Cl_2 and b) CH_2Cl_2 with 1 % (w/v) NH_4OH containing 0.1 M TBAP.

Table 1. UV/Visible spectroscopic data of investigated cobalt corroles under different solution conditions.

Solution condition	Cpd	Soret region		Wavelength: λ_{max} , nm ($\epsilon \times 10^{-4} \text{ M}^{-1} \text{ cm}^{-1}$)			
				visible region			
CH_2Cl_2	1DMSO	379 (7.2)	416 (4.1) sh		566 (1.3)		
	2DMSO	385 (6.4)			562 (1.3)		
	3DMSO	387 (7.1)	414 (3.6) sh		563 (1.3)		
	4DMSO	382 (7.2)			563 (1.6)		
CH_2Cl_2 +1 % NH_4OH	1NH₃		417 (5.0)	438 (8.3)	451 (7.6)	586 (1.9)	615 (2.2)
	2NH₃		407 (2.3)	436 (4.1)	457 (3.9)	541 (0.6)	621 (1.7)
	3NH₃		420 (3.9)	434 (4.9)	453 (4.2)	526 (1.1)	628 (2.1)
	4NH₃		400 (4.2)	433 (5.6)	452 (5.2)	581 (1.4)	623 (2.4)

the most intense of which is located at 615–628 nm and is considered as a marker band for the hexacoordinate Co^{II} corrole. The overall spectral patterns of the **1NH₃** to **4NH₃** derivatives are similar to spectral patterns earlier reported for structurally related cobalt corroles containing two axially bound pyridine ligands.^[1]

The same UV/Vis spectrum is obtained for **1NH₃** in CH_2Cl_2 with and without 1 % NH_4OH , both spectra being essentially identical with spectra of the initial synthesized bis- NH_3 adduct in toluene (see Exp. Sect.). In contrast, different spectral patterns are seen upon going from **2NH₃** to **4NH₃** in CH_2Cl_2 with and without 1 % NH_4OH . The largest variation between the two solution conditions is observed in the Soret region of the spectrum, where bands assigned to a five-coordinate corrole are observed for the bis- NH_3 derivatives **2**, **3** and **4** but not for compound **1**. The fact that identical spectra are obtained for **1NH₃** in CH_2Cl_2 with and without added NH_4OH suggests that this corrole has a stronger affinity for NH_3 than the other investigated derivatives and that dissociation of the bound NH_3 axial ligands does not readily occur in CH_2Cl_2 . This observation is consistent with the lower electron density of the **1NH₃** macrocycle due to the electron-withdrawing halogen atoms on the *meso*-phenyl rings of the corrole and it is also consistent with the relative order of the pyridine binding constants to the mono-DMSO derivatives, as described below.

Pyridine Binding Ability

The relative intensity of the diagnostic six-coordinate absorption band at 615–623 nm for compounds **1NH₃** to **4NH₃** in Figure 3a follows the order **1NH₃** > **2NH₃** > **3NH₃** > **4NH₃** and provides an approximate measure of the extent of NH_3 dissociation from the bis- NH_3 adducts in CH_2Cl_2 at a corrole concentration of 10^{-3} M. The band intensity follows the order **1** > **2** > **3** > **4** and this would also be the relative order of the NH_3 binding strength for the examined bis- NH_3 derivatives in the CH_2Cl_2 solvent.

We could not evaluate the magnitude of the NH_3 binding constants to the four-coordinate corroles in the current study but we could measure binding constants for the stepwise addition of two pyridine molecules to the mono-DMSO adducts in CH_2Cl_2 . This was done by monitoring changes in the UV/Visible spectra during a titration of each corrole with pyridine, giving

results of the type shown in Figure 4 for the stepwise conversion of **1DMSO**–**4DMSO** to their respective six-coordinate, bis-pyridine form. Equilibrium constants for addition of the first pyridine axial ligand to the mono-DMSO adducts were not determined due to the small differences in spectra between the initial corrole reactant and the mono-pyridine adduct but much larger spectral changes and well-defined log-log plots were obtained upon formation of the bis-pyridine derivatives as shown in Figure 4, b. The binding constant for addition of a second pyridine molecule to **1DMSO** was measured as $\log K_2 = 3.6$ and this $\log K_2$ value systematically decreased to 2.1, 1.5 and 1.1 for **2DMSO**, **3DMSO** and **4DMSO**, respectively. These values are given in Table 2 along with $\log K_2 = 2.6$ for addition of a second pyridine molecule to $\text{Mes}_3\text{CorCo}(\text{py})$, as reported in an earlier study.^[1]

MS Spectra of **1NH₃**–**4NH₃** and **1DMSO**–**4DMSO**

Mass spectrometry utilizing different ionization modes (ESI and MALDI-TOF) was used to characterize the bis-ammonia and mono-DMSO complexes. In the case of MALDI-TOF analysis, the data gives in each case the cobalt corroles without any axial ligands, the DMSO and NH_3 groups being lost under high vacuum during the preparation step, thus leading to the tetra-coordinated cobalt complexes **1**–**4** (Figure S2a–h). In order to observe the labile ligands, we optimized the MS procedure in the ESI mode by directly injecting the compounds dissolved in a solution of ammonia or DMSO (see Exp. Sect.) and, under these conditions, the labile DMSO or NH_3 axial ligands were both observed (see Figure S2a–h) providing further evidence for formation of the penta- and hexacoordinate cobalt corrole complexes in the original synthesis.

¹H NMR, EPR and IR Characterization

The effect of the *meso*-aryl substituents on the coordinating ability of the cobalt center of **1NH₃** and **4NH₃** was examined by ¹H NMR spectroscopy. In the polar CDCl_3 solvent, broad signals were obtained for **1NH₃** which contains two electron-withdrawing Cl substituents on the *meso*-phenyl rings (Figure 5c), but almost no signal is observed for **4NH₃** which bears electron-donating substituents (Figure 5f). The lack of a diamag-

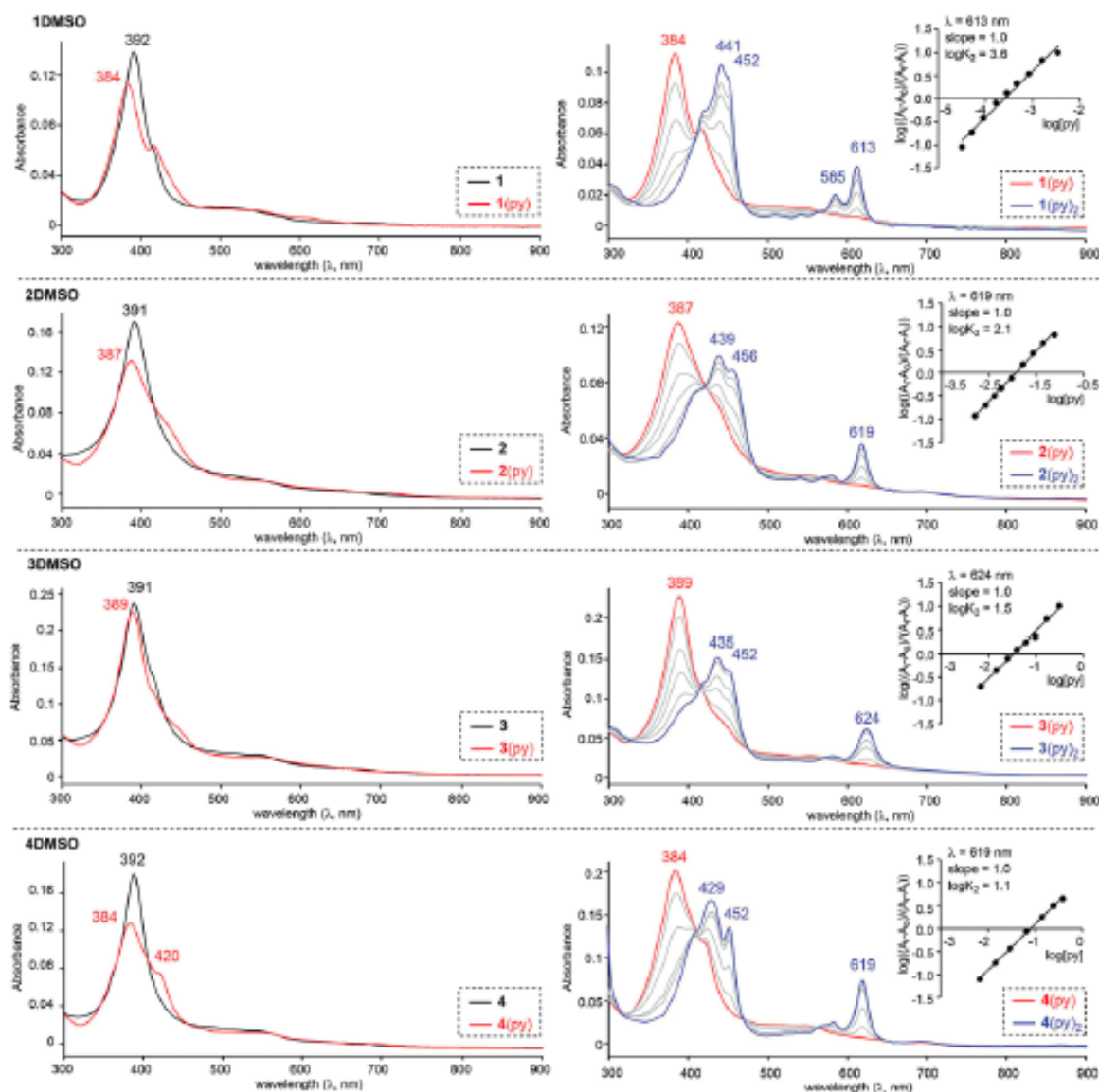


Figure 4. UV/Visible spectra changes of 1DMSO-4DMSO (at $\approx 10^{-5}$ M) in CH_2Cl_2 upon addition of pyridine to solution.

Table 2. Pyridine binding constants, $\log K$, for cobalt corroles in CH_2Cl_2 .

Compound	$\log K_1$	$\log K_2$	Reference
(<i>o,o</i> - Cl_2Ph) ₂ CorCo 1	–	3.6	tw ^[a]
(<i>p</i> -COOMePh)(Mes) ₂ CorCo 2	–	2.1	tw
(Ph) ₂ CorCo 3	–	1.5	tw
[<i>o,o</i> - <i>p</i> -(OMe) ₂ Ph](Mes) ₂ CorCo 4	–	1.1	tw
(<i>m</i> -NO ₂ - <i>p</i> -FPh)(<i>p</i> -NO ₂ Ph) ₂ CorCo	–	3.9	[8]
(<i>m</i> -NO ₂ - <i>p</i> -OMePh)(<i>p</i> -NO ₂ Ph) ₂ CorCo	–	3.8	[8]
(<i>m</i> -NO ₂ - <i>p</i> -FPh)(Mes) ₂ CorCo	–	3.3	[8]
(<i>m</i> -NO ₂ - <i>p</i> -OMePh)(Mes) ₂ CorCo	–	3.0	[8]
(Mes) ₂ CorCo	6.5	2.6	[8]

[a] tw = this work.

netic spectrum for 4NH_2 is consistent with a dissociation of one or both labile NH_2 ligands in CDCl_3 , generating a paramagnetic form of the cobalt corrole due to the non-innocent character of the four- or five-coordinate species.^[6] However, when NH_3 gas was bubbled into the NMR tube, broad signals appeared for 4NH_3 (Figure 5e) indicating a mixture of electronic spin states for the cobalt ion ($S = 0$, diamagnetic and $S = 1$, paramagnetic), as well as some $S = 1/2$ species probably due to traces of the corrole radical.^[5c] This contrasts with the NMR spectrum of 1NH_2 , which exhibits sharp signals (Figure 5b) characteristic of a pure diamagnetic compound.

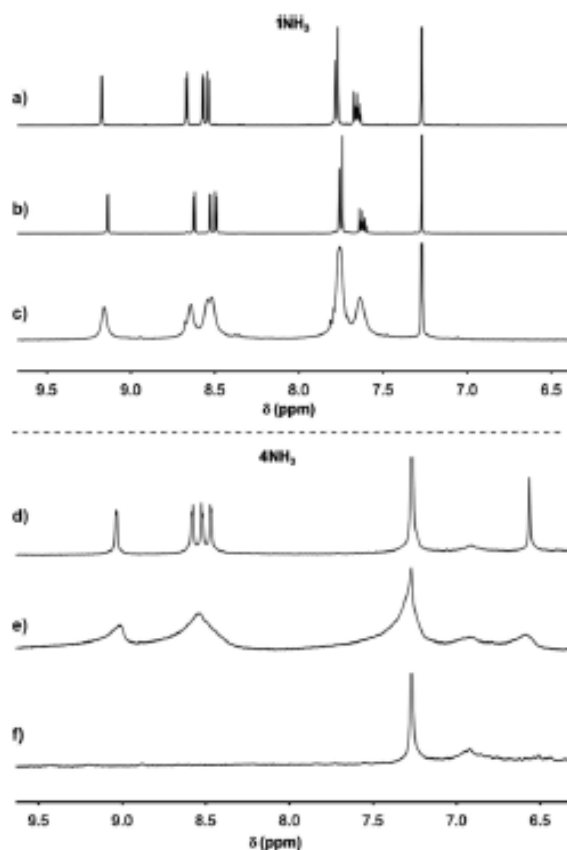


Figure 5. ^1H NMR spectra of 1NH_2 (top) and 4NH_2 (bottom) in CDCl_3 before (c) and (f) and after addition of (b), (e) $\text{NH}_3(\text{g})$ or (a), (d) $\text{N}_2\text{H}_4\cdot\text{H}_2\text{O}$.

In order to confirm the paramagnetic behavior of these complexes, EPR experiments were run on the electron-poor and electron-rich 1NH_2 and 4NH_2 complexes in toluene at 293 K (Figures S9 and S10), and this data should give a direct measurement of the electronic spin state.^[6] In both cases, the spectra (Figures S9a and S10a) have a radical-like pattern centered at $g = 2.002$, with a small broadening of the single line (FWHH = 10 G). This central line can be assigned to a radical impurity due to over-oxidation of the corrole that is frequently observed for these type of macrocycles.^[1c,5b]

No other transition is shown for 1NH_2 , indicating either that the cobalt corrole is mainly in a diamagnetic Co^{II} form, or that it possesses a strongly antiferromagnetic $S = 0$ spin state and a Co^{I} -corrole $^{2-}$ formulation. The assignment is favored after addition of NH_4OH , since two broad signals with low intensities assigned to the excited triplet state now appear on either side of the radical signal. Under these conditions, the hexacoordinate complex with two NH_2 axial ligands is predominant in solution (see UV/Visible data). It should also be noted that the intensity of the radical signal strongly decreases after addition of NH_4OH , suggesting on one hand the formation of a more diamagnetic hexacoordinate complex, and on the other a partial over-oxidation of the pentacoordinate corrole formed after loss of one NH_2 ligand from the bis-ammonia complex when it is dissolved in toluene (Figure S9a). In addition, the low intensity of the triplet state agrees with a large singlet-triplet gap, inducing a weak population of the excited triplet state. This result adds strong support to assignments from the NMR spectra which show formation of a diamagnetic 1NH_2 complex dissolved in solutions containing NH_4OH or N_2H_4 . Finally, it should be noted that hydrazine can act as a mild reducing agent which would remove all traces of any paramagnetic corrole radical.

A rather similar situation was observed for 4NH_2 (Figure S10a,b) which is characterized by a radical spectrum centered at $g = 2.002$, but flanked with two strong signals at $g = 2.015$ and 1.986 assigned to the triplet state resulting from ferromagnetic coupling between the cobalt(II) ion and the ion radical for the Co^{I} -corrole $^{2-}$ formalism. In addition, the spectrum is devoid of any resolvable hyperfine interactions due to the strongly delocalized spin system. As the intensity of the spectrum is directly correlated to the relative population of the triplet state, one can anticipate a lower singlet-triplet gap compared to 1NH_2 and an overall picture of a singlet ground state with the thermally accessible triplet state.

The addition of N_2H_4 (Figure S10c) resulted in a loss of the radical signal and only the two signals assigned to the triplet state are preserved. This result is in agreement with the NMR spectroscopic data of 4NH_2 that show more resolved signals after addition of hydrazine (Figure 5d). The rather broad NMR lines of 4NH_2 compared to 1NH_2 can also be attributed to a residual paramagnetism of the complex due to the more accessible triplet state.

Moreover, the hexacoordinate bis- NH_2 complexes could be formed directly from the pentacoordinate mono-DMSO derivatives by adding ammonia gas into the NMR tube in order to displace the coordinated DMSO. After the displacement has occurred, a peak for one free DMSO molecule is seen in the spectrum at $\delta = 2.60$ ppm and the two NH_2 ligands appear in the anisotropic ring current of the corrole at around -6.60 . Integration of the NMR signals for both DMSO and NH_2 peaks confirms the assigned penta- and hexa-coordination (see Figure S3e-g).

The axial coordination of NH_2 was also observable by infrared spectroscopy. The infrared spectra of 1NH_2 - 4NH_2 are characterized by antisymmetric and symmetric NH_2 stretching vibrations at around 3260 and 3360 ppm, respectively (see Figure S6 and Table S2).

Penta- vs. Hexa- Coordination of Cobalt Corroles Examined by DFT Computational Study

Cobalt corroles are known to bind two molecules of pyridine but only a single molecule of carbon monoxide^[27] or DMSO.^[28] In the current study, we observed that the examined cobalt corroles would easily bind two molecules of NH_3 and this was theoretically examined for the case of $(\text{Cl}_2\text{Ph})_3\text{CorCo } \mathbf{1}$, which was also modelled with respect to the binding of CO, DMSO, pyridine and NH_3 . The geometry of the cobalt corrole $(\text{Cl}_2\text{Ph})_3\text{CorCo } \mathbf{1}$ was optimized, both as the four-coordinate complex and as five- and six-coordinate derivatives complexed with one and two axial ligands, respectively. The axial coordination processes are described by Equation (1) and Equation (2), where X represents the axial ligand and $\Delta_r G$ is the complexation free energy. The calculated complexation free energies are listed in Table 3.



A negative value of $\Delta_r G_2(\text{X})$ indicates that hexa-coordination is favored, while a positive value indicates complexation of only a single axial ligand at the cobalt center. As seen in Table 3, the formation of hexacoordinate complexes for $(\text{Cl}_2\text{Ph})_3\text{CorCo } \mathbf{1}$ is favorable with NH_3 or pyridine ligands, but unfavorable with both CO and DMSO, a result in agreement with the experimental data.^[27] However, it should be noted that the theoretical value of $\Delta_r G_2(\text{DMSO})$ is close to 0, a result which may explain why the characteristic UV/Visible spectral pattern of a bis-DMSO complex is observed for **1DMSO** in the DMSO solvent at a 10^{-3} M concentration of the corrole (Figure 2b).

In order to theoretically study the effect of different meso-phenyl substituents on the NH_3 binding ability of the corroles, we also calculated the complexation free energies of **1–4** with NH_3 . These values are given in Table 4 and show that the bis-ammonia complex is in each case greatly favored, with large negative values ranging from -12.2 to -9.3 kcal/mol. The data in Table 4 show that the binding strength with NH_3 follows the

order: **1** > **2** > **4** > **3**, which is slightly different from the order predicted by measurements of the six-coordinate marker band in UV/Visible spectra for the same compounds, namely **1** > **2** > **3** > **4**.

Single Crystal X-ray Diffraction Studies of Bis-Ammonia Cobalt Corrole **1NH₂**

Single crystals of corrole **1NH₂** were grown by slow evaporation of a CH_2Cl_2 /hexane mixture of **1NH₂** and gave the structure shown in Figure 6. Details about the structure are given in the SI.

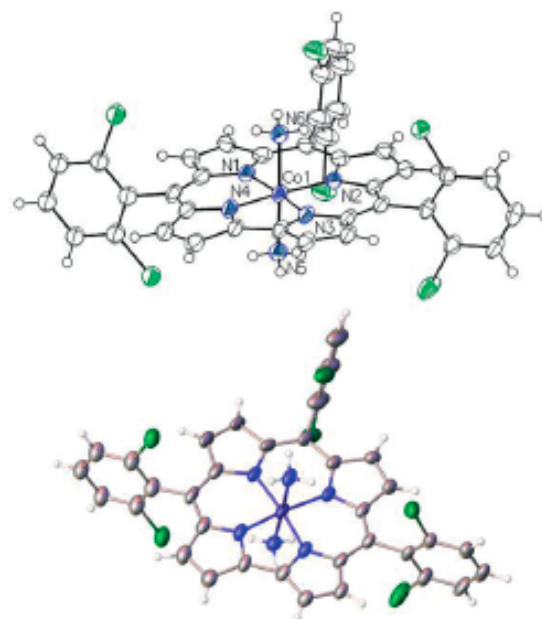


Figure 6. ORTEP view of compound **1NH₂**. Thermal ellipsoids are drawn at the 50% probability level. Solvent and disordered part are omitted for clarity.

Table 3. Complexation free energies $\Delta_r G_1(\text{X})$ and $\Delta_r G_2(\text{X})$ [kcal/mol] of $(\text{Cl}_2\text{Ph})_3\text{CorCo } \mathbf{1}$ with NH_3 , pyridine, CO and DMSO.

Cobalt corrole	Ligand X	Complexation free energies [kcal/mol]	
		$\Delta_r G_1(\text{X})$	$\Delta_r G_2(\text{X})$
1NH₂	NH_3	-4.7	-12.2
1Py	Pyridine	-9.0	-11.5
1CO	CO	-2.3	+9.4
1DMSO	DMSO ^[a]	-3.3	+1.4

[a] S = 1 is the most stable state for these complexes.

Table 4. Complexation free energies $\Delta_r G_1(\text{X})$ and $\Delta_r G_2(\text{X})$ [kcal/mol] of cobalt corroles **1–4** with NH_3 .

Compound	Ligand X	Complexation free energies [kcal/mol]	
		$\Delta_r G_1(\text{X})$	$\Delta_r G_2(\text{X})$
$(\text{Cl}_2\text{Ph})_3\text{CorCo } \mathbf{1}$	NH_3	-4.7	-12.2
$(\text{COOMePh})\text{Mes}_2\text{CorCo } \mathbf{2}$	NH_3	-3.1	-10.0
$(\text{Ph})_3\text{CorCo } \mathbf{3}$	NH_3	-4.5	-9.3
$[(\text{OMe})_2\text{Ph}]\text{Mes}_2\text{CorCo } \mathbf{4}$	NH_3	-2.9	-9.7

Compound **1NH₂** crystallizes in a *P2₁/c* centrosymmetric group. The asymmetric unit is composed of a cobalt corrole with two disordered dichlorophenyl units over two positions and one disordered hexane molecule (see SI). The cobalt complex is hexacoordinate with four nitrogen atoms of the corrole unit and two axial ammonia ligands which are bound to the cobalt center in an apical position. Thus, this complex adopts an octahedral geometry with a typical *Oh* type symmetry.

The Co–N_{1,2,3,4-corrole} distances range from 1.868(6) to 1.908(6) Å, values almost identical to those reported for the structure of vitamin B12,^[4] i.e. between 1.887 Å and 1.912 Å. The axial Co–NH₃ distances of 1.964(6) Å and 1.974(6) Å are slightly longer than the Co–N_{1,2,3,4-corrole} distances. However, the values are still of the same order of magnitude as in the Co^{III}–NH₃ complex 1.961(20) Å, and are shorter than the cobalt–pyridine distances for the same type of compound, 1.981–1.990(2) Å.^[9] The corrole complex has a planar geometry with a maximum deviation from the mean plane of 0.13 Å for the C36 atom. The crystalline stacking is governed by intermolecular hydrogen bonds between the NH₃ and the Cl groups of the corrole (*ortho*-dichlorophenyl substituents), thus forming a molecular wire bonded by hydrogen bonds (Figure S2).

Electrochemistry

Redox properties of the synthesized corroles were investigated by cyclic voltammetry under different solution conditions. As previously demonstrated for a related series of derivatives,^[1] the synthesized mono-ligated DMSO cobalt corroles are invariably characterized by facile and reversible one-electron reductions in CH₂Cl₂ or DMSO (Figure 7 and Figure S7). This contrasts with what is seen for the hexacoordinate bis-pyridine derivatives where the first reduction in pyridine is irreversible and occurs at much more negative potentials.^[2] The difference in redox behavior between what is seen in CH₂Cl₂, DMSO or pyridine was explained as being due to a different site of electron transfer and a different electronic configuration of the redox active form of the corrole in solution, the mono-DMSO adducts being assigned as containing a Co^I center and a non-innocent easily reduced macrocyclic ligand in CH₂Cl₂ or DMSO as compared to a Co^{III} center and a hard to reduce innocent macrocyclic ligand in pyridine where the first reduction involves a metal-centered Co^{III}/Co^{II} process followed by a loss of one or both axial ligands after electron transfer. In other words, the reversibility of the first reduction process and the small electronic effect of the ligand on the potential values are in accordance with a Co^{II} center and a non-innocent macrocycle.

The difference in axial coordination and ligand innocence or non-innocence also manifests itself in a difference of half-wave potentials for the first one-electron oxidation, this process being easiest (more negative potential) for oxidation of the cobalt bis-pyridine derivatives in pyridine and harder (more positive potential) for oxidation of the four- or five-coordinate cobalt corroles in CH₂Cl₂ or DMSO.^[1]

We anticipated observing facile one-electron reductions for the newly synthesized mono-DMSO derivatives in the current study and this is indeed the case as shown by the cyclic voltammograms measured in DMSO (Figure 7) or CH₂Cl₂ (Figure S7).

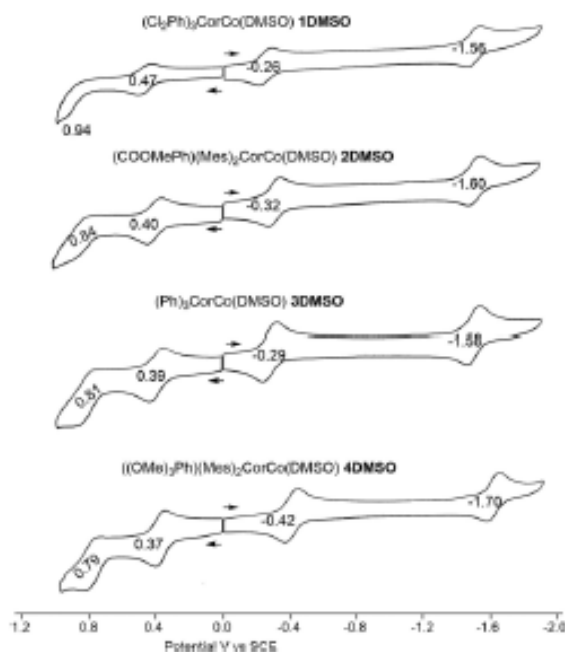


Figure 7. Cyclic voltammograms of cobalt corrole complexes **1DMSO**–**4DMSO** ($\approx 10^{-2}$ M) in DMSO containing 0.1 M TEAP.

1DMSO–**4DMSO** in DMSO (Figure 7) undergoes two reductions and two oxidations within the negative and positive potential limits of the solvent. The first one-electron reduction is located at $E_{1/2} = -0.26$ to -0.42 V and the second at -1.56 to -1.70 V. The first one-electron oxidation ranges from $E_{1/2} = 0.37$ to 0.47 V depending upon the specific *meso*-substituents. In addition, the absolute potential difference between the first reversible reduction and the first reversible oxidation of each compound (the HOMO–LUMO gap) varied from 0.72 to 0.79 V in DMSO, a range of values slightly smaller than the measured HOMO–LUMO gaps in CH₂Cl₂, where the potential separation ranges from 0.75 to 0.87 V (See examples of cyclic voltammograms in Figure S7 and summary of half-wave potentials in Table 5).

The binding of NH₃ to the neutral and singly oxidized cobalt corroles **1** to **4** is stronger than the binding of DMSO to the same series of compounds in CH₂Cl₂ and this is reflected in large negative shifts of potential for the first one-electron abstraction, the measured $E_{1/2}$ values ranging from 0.30 V for **1NH₂** to 0.05 V for **4NH₂** as seen in Figure 8 (and Figure S8). Like potentials for electron abstraction from the mono-DMSO adducts, the measured half-wave potentials for oxidation of the bis-NH₃ derivatives shift negatively with increasing electron-donating character of substituents on the three *meso*-phenyl rings of the macrocycle. However, it should be noted that the substituent effect of the *meso*-phenyl groups is larger for the bis-NH₃ ligated Co^{III} corroles than for the five-coordinate derivatives (perhaps containing a Co^I center and a non-innocent ligand), resulting in an increasing difference in $E_{1/2}$ values between the first oxidation of the same corrole with different axial

Table 5. Half-wave or peak potentials ($E_{1/2}$ or E_p , V vs. SCE) for reductions and the first two oxidations of the cobalt corroles in different solution conditions containing 0.1 M TBAP (sweep rate: 100 mV/s).

Solution condition	Cpd	oxidation	Potentials (V vs. SCE)			
			reduction		oxidation	
CH_2Cl_2	1DMSO		0.99	0.71	-0.15	-1.62 ^[a]
	2DMSO		0.86	0.62	-0.20	-1.68 ^[a]
	3DMSO	11.48	0.83	0.64	-0.11	-1.62 ^[a]
	4DMSO	11.45	0.88	0.60	-0.27	-1.78 ^[a]
DMSO	1DMSO		0.94 ^[a]	0.47	-0.26	-1.56
	2DMSO		0.84	0.40	-0.32	-1.60
	3DMSO		0.81	0.39	-0.29	-1.58
	4DMSO		0.79	0.37	-0.42	-1.70
CH_2Cl_2	1NH ₃	11.74	0.92	0.30	-0.07	-1.08
	2NH ₃	11.49	0.74	0.14	-0.19	-1.29
	3NH ₃	11.44	0.68	0.11	-0.14 ^[a]	-1.21
	4NH ₃	11.45	0.68	0.05	-0.26	-1.40
CH_2Cl_2 (w/v 1 % NH ₄ OH)	1NH ₃		-	-	-	-
	2NH ₃		0.74	0.14	-0.19	-1.66 ^[a]
	3NH ₃		0.64	0.10	-0.23 ^[a]	-1.64 ^[a]
	4NH ₃		0.69	0.05	-0.26	-1.78 ^[a]

ligands, from 170 mV in the case of compound **1** (0.47 V in DMSO and 0.30 V in CH_2Cl_2) to 320 mV for compound **4** (0.37 vs. 0.05 V).

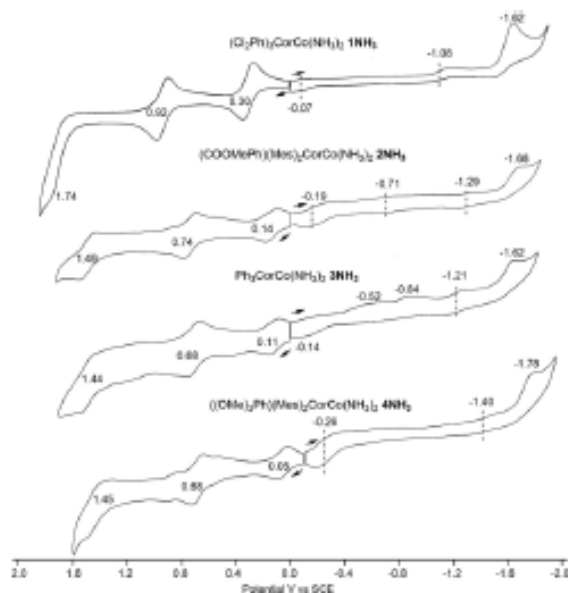


Figure 8. Cyclic voltammograms of cobalt corrole complexes **1NH₃**–**4NH₃** ($\times 10^{-2}$ M) in CH_2Cl_2 containing 0.1 M TBAP.

The oxidation of $(\text{Cor})\text{Co}(\text{NH}_3)_2$ to $[(\text{Cor})\text{Co}(\text{NH}_3)_2]^+$ occurs in the absence of coupled chemical reactions but this is not the case for the reduction where a dissociation of one or both NH_3 axial ligands is proposed to occur prior to electron transfer, thus leading to several competing reduction pathways, one for the six-coordinate complex which is the most difficult to reduce and located at $E_{1/2}$ values more negative than -1.0 V vs. SCE and

the others for the five- and/or four-coordinate species which generate the Co^{II} corrole at potentials as positive as -0.07 to -0.26 V as seen in Figure 8.

Evidence for dissociation of the NH_3 ligand(s) prior to electron transfer is given in part by the shape of the current voltage curves which is consistent with a coupled chemical reaction prior to electron transfer (an electrochemical CE mechanism), in part by the fact that the first (ill-defined) reduction potentials of the bis- NH_3 complexes in Figure 8 are almost identical to the well-defined $E_{1/2}$ values for the same cobalt corroles bearing a single DMSO axial ligand in their synthesized air stable form (Figure 7), namely, -0.19/-0.20 V for **2NH₃** and **2DMSO**, -0.11 V for **3NH₃** and **3DMSO** and -0.26/-0.27 V for **4NH₃** and **4DMSO** (see Figure 7 and Table 5), and in part by the current voltage curves for **1NH₃** to **4NH₃** obtained in CH_2Cl_2 containing excess NH_3 added in the form of NH_4OH .

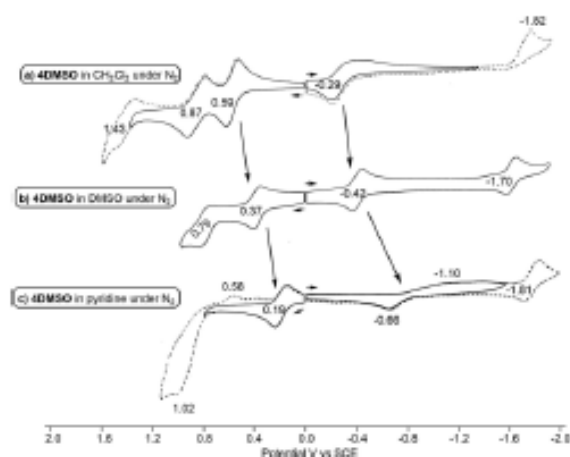


Figure 9. Cyclic voltammograms of **4DMSO** in a) CH_2Cl_2 , b) DMSO and c) pyridine containing 0.1 M TBAP.

Examples of cyclic voltammograms under these later solution conditions are given in Figure S8 for **2NH₃**, **3NH₃**, and **4NH₃**. For all three corroles, the half-wave potentials of the two oxidations are the same as in the CH₂Cl₂ solvent without added NH₃ (Figure 8), consistent with the proposed [(Cor)Co(NH₃)₂]^{0/+} redox process. However, when NH₃ is added to solution, a diminished current intensity is seen in the first reduction peak as the equilibrium is shifted towards the harder to reduce bis-ligated form of the corrole in solution.

As mentioned above, the average HOMO–LUMO gap of the four mono-DMSO adducts **1DMSO** to **4DMSO** decreases from 825 mV in CH₂Cl₂ (Figure S7) to 705 mV in DMSO (Figure 7) and the gap decreases even further for the bis-NH₃ adducts where the separation between the first oxidation and first reduction is 330 mV for **2NH₃** in CH₂Cl₂ and 310 mV for **4NH₃** under the

same solution conditions (see Figure 8). In each case, the decrease in the measured gap is attributed for the most part to a shift of the HOMO towards an easier oxidation. As shown in Figure 9, the measured $E_{1/2}$ value for the one-electron abstraction of **4DMSO** varies from 0.59 to 0.37 V upon going from CH₂Cl₂ to DMSO and from 0.37 to 0.19 V upon going from DMSO to pyridine.

Spectroelectrochemistry

The singly oxidized and singly reduced forms of mono-DMSO and bis-ammonia corroles were characterized by thin-layer UV/Visible spectroelectrochemistry and examples of the resulting spectral changes are illustrated in Figure 10 for compounds **4DMSO** and **4NH₃** under different solution conditions. The

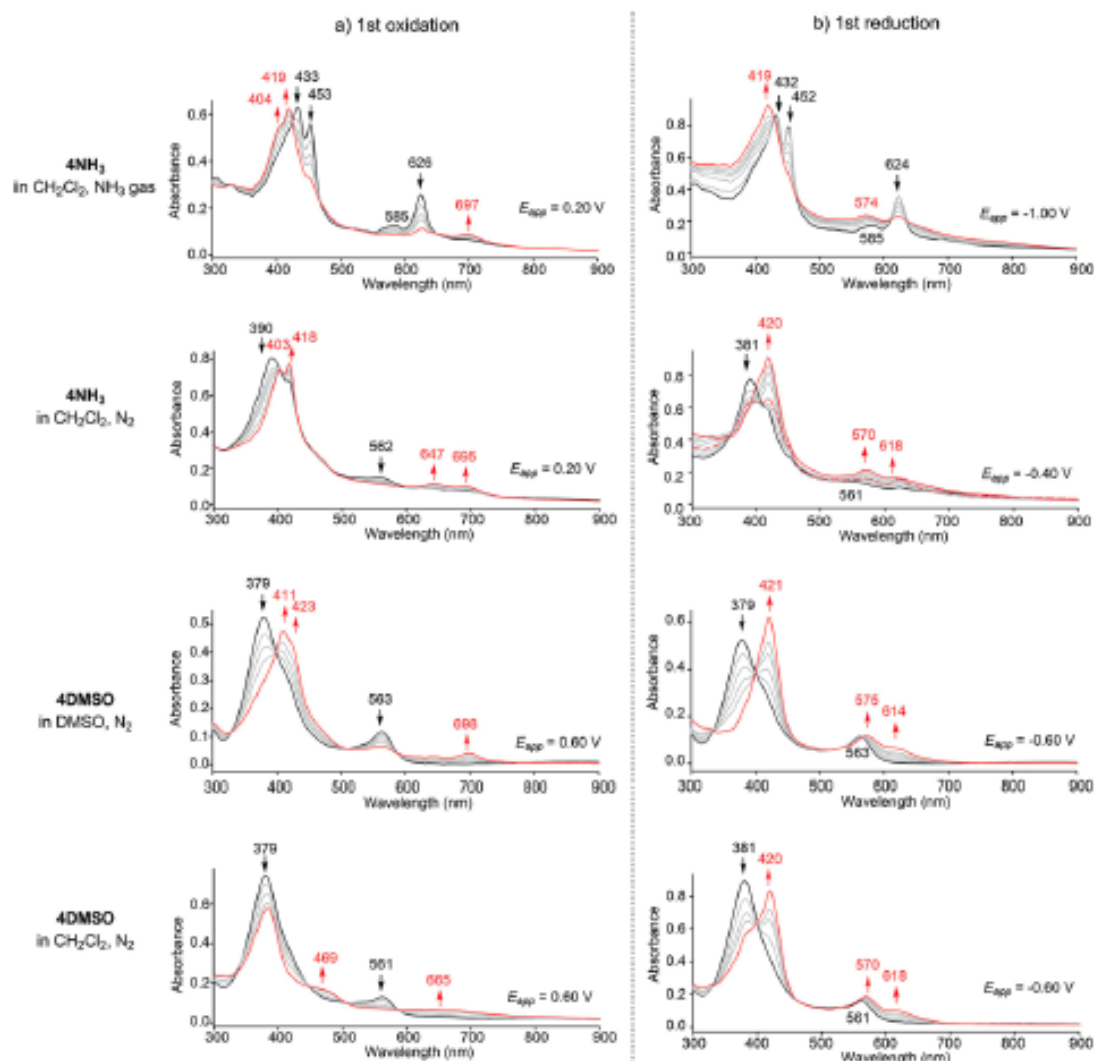


Figure 10. UV/Visible spectra of a) the first oxidations and b) the first reductions for **4DMSO** under controlled potentials in different solution condition containing 0.1 M TBAP.

spectral changes obtained during reduction of **4DMSO** are essentially independent of solvent as shown in Figure 10b where the Soret band of the neutral corrole shifts from 379/381 to 420/421 nm and the single visible band at 561/563 nm of the unreduced compound is replaced by two visible bands located between 570 and 618 nm after the addition of one-electron.

The spectrum of **[4DMSO]⁻** is assigned as containing an unligated Co^{II} central metal ion and virtually the same UV/Visible spectrum is obtained after the reduction of **4NH₂** in CH₂Cl₂ under an N₂ or NH₃ atmosphere, thus suggesting a similar square planar Co^{II} electroreduction product. However, as mentioned earlier in the manuscript, a partial dissociation of one or both NH₂ ligands occurs for **4NH₂** when dissolved in CH₂Cl₂ under N₂ but the original bis-NH₂ adduct can be easily regenerated under an NH₃ atmosphere or when NH₄OH is added to solution. This is shown by the spectroscopic data in Figure 10 where the initial spectrum of **4NH₂** under N₂ is identical to that of four- or five-coordinate **4DMSO** species but converts to that of the original bis-NH₂ adduct when NH₃ gas is bubbled through the solution. Under these conditions, the neutral six-coordinate complex is characterized by a split Soret band at 432 and 452 nm and an intense six-coordinate marker band at 624 nm as shown in Figure 10. This spectrum is assigned to the Co^{II} corrole which is then converted to its Co^{III} form after controlled potential reduction at -1.00 V in the thin layer cell. The same spectral product is obtained after the one-electron oxidation of **4NH₂** under N₂ or NH₃ and this species, with a split Soret band between 403 and 419 nm and a broad Q band between 600 and 725 nm is assigned as a six-coordinate Co^{III} π -cation radical. Singly oxidized **4DMSO** has a similar spectrum in DMSO which is assigned to the bis-DMSO Co^{III} π -cation radical but this spectrum differs substantially from the spectrum of singly oxidized **4DMSO** in CH₂Cl₂.

Conclusions

In summary, we have reported and characterized the electrochemistry and spectroscopic properties of mono-DMSO and bis-ammonia cobalt corroles in CH₂Cl₂, DMSO and CH₂Cl₂/NH₄OH mixtures. The ligated DMSO axial ligand can be replaced by two ammonia molecules, shifting from a pentacoordinate cobalt corrole to a hexacoordinate one. The two ammonia molecules can be removed to smoothly access the tetra-coordinated cobalt corrole, the active form of the compound able to selectively bind carbon monoxide. The formation of pentacoordinate cobalt corroles in the case of the DMSO and CO derivatives and hexacoordinate derivatives in the case of NH₃ and pyridine axial ligands is observed experimentally and also observed by a DFT computational study. Thus, the theoretical predictions are in agreement with results from an X-ray determination on one compound and a variety of experimental data on all of the examined derivatives, which were characterized by electrochemistry, UV/Visible and ¹H NMR, spectroscopy and mass spectrometry.

Experimental Section

Materials and Instrumentation: All chemicals and solvents were of the highest grade available and were used without further purifi-

cation. Dichloromethane (CH₂Cl₂) and pyridine (Py) were purchased from Sigma Aldrich Co. and used as received. Tetra-*n*-butyl-ammonium perchlorate (TBAP), used as supporting electrolyte, was obtained from Sigma Chemical Co.

Silica gels 60 (70–230 and 230–400 mesh, Sigma Aldrich) were used for column chromatography. Reactions were monitored by thin-layer chromatography, UV/Visible spectroscopy and mass spectrometry. CDCl₃ was used as a solvent for ¹H NMR spectra which were recorded on a Bruker AVANCE III spectrometer (500 MHz). The measurements were made at the PACSMUB-WPCM technological platform, which relies on the "Institut de Chimie Moléculaire de l'Université de Bourgogne" and Wellence "TM", a Burgundy University private subsidiary. Chemical shifts were expressed in ppm. For **1DMSO-4DMSO**, NH₃ or N₂H₄·H₂O was added to enhance resolution of the spectra. Mass spectra were recorded on a Bruker Ultraflex Extreme MALDI Tandem TOF Mass Spectrometer using diethanol as the matrix or on a LTQ Orbitrap XL (Thermo) Instrument in the ESI mode (for the HRMS spectra). Low-resolution mass spectra (LRMS) were recorded on a Thermo Scientific™ MSQ Plus™ single quadrupole mass spectrometer equipped with an electrospray ionization (ESI) source. All LRMS mass spectra were recorded in the positive mode (ESI⁺) using the following parameters: "full scan", 250–1250 a.m.u., data type: centroid, needle voltage: 3.0 kV, detector voltage: 1100 V, probe temperature: 350 °C, cone voltage: 10 V or 75 V and scan time: 1 s. Sample preparation: Stock solution (1 mg/mL) of cobalt corroles were prepared in CH₃CN. Dilution by a factor $\times 20$ was achieved in CH₃CN/H₂O (2:1, v/v) and the resulting solution (ca. 600 μ L) was directly analyzed by ESI-MS (direct infusion mode with a syringe pump, cone voltage: 75 V). To assess the chelation of DMSO and NH₃ molecules, 2 μ L of an aq. DMSO solution (25 % in water, v/v) or NH₄OH was added to a dilute sample prior to injection through a direct infusion mode (cone voltage: 10 V).

Infrared spectra were recorded in the solid state on an FT IR Bruker Vertex 70v, using KBr pellets. UV/Visible spectra were recorded with a Hewlett-Packard Model 8453 diode array spectrophotometer or with a Varian Cary 50. Cyclic voltammetry was carried out at 298 K at 100 mV/s sweep rate using an EG&G Princeton Applied Research 173 potentiostat/galvanostat. A three-electrode system was used for cyclic voltammetric measurements and consisted of a glassy carbon working electrode, a platinum counter electrode and a saturated calomel reference electrode (SCE). The SCE was separated from the bulk of the solution by a fritted glass bridge of low porosity which contained the solvent/supporting electrolyte mixture. Thin-layer UV/Visible spectroelectrochemical experiments were performed with a commercial spectroelectrochemical cell purchased from Pine Instrument Inc. Potentials were applied and monitored with an EG&G PAR Model 173 potentiostat. High purity N₂ from Trigas was used to deoxygenate the solution and a stream of nitrogen was kept over the solution during each electrochemical and spectroelectrochemical experiment. Thermogravimetric (TGA) analyses were recorded using a Netzsch STA 409 PC thermal analyzer. The samples (approximately 7.0 mg) were heated from 298 to 1273 K with a heating rate of 5 K min⁻¹ under a flow of nitrogen (30 mL min⁻¹) and oxygen (10 mL min⁻¹).

DFT Computational Details: Quantum mechanics calculations were performed with the Gaussian09 software package.^[10] Energy and forces were computed by density functional theory with the hybrid B3LYP exchange-correlation functional. Dispersion effects are included through the empirical correction of Grimme with the Becke–Johnson damping scheme.^[11] Previous studies have shown that this level accurately reproduces geometries and energies of cobalt corrole complexes.^[14,12]

The solvent effects were modelled using a polarizable continuum model^[13] (PCM) as implemented in Gaussian09 to describe the dichloromethane bulk medium. Geometries were optimized and characterized with the 6-31+G(d, p) basis sets for all atoms. Frequency calculations were performed to ensure the absence of any imaginary frequencies on local minima, and the presence of only one imaginary frequency on transition states. Gibbs free energies were computed at 298 K and 1 atm. using the electronic energies and frequencies computed at the 6-31+G(d, p) level.

Synthesis of Free-base and Mono-DMSO Cobalt Corroles. General Procedure: The free-base corroles tris(2,6-dichlorophenyl) corrole (**1H**), 10-(4-methoxycarbonylphenyl)-5,15-dimesitylcorrole (**2H**), triphenylcorrole (**3H**), 10-(2,4,6-trimethoxyphenyl)-5,15-dimesitylcorrole (**4H**) were synthesized according to procedures reported by Gryko and co-workers.^[2] Each free-base corrole, **1H**, **2H**, **3H** or **4H**, (1.0 equiv.) was added to a solution of cobalt acetate tetrahydrate (47 mg, 189 μ mol, 1.2 equiv.) with DMSO (20 mL) in a round-bottomed flask after which the reaction mixture was stirred at 80 °C for 40 minutes and then cooled to room temperature. The crude mixture, **Crude-Co**, was poured into a cold NaCl aqueous solution (0.8 M), the resulting suspension was filtered, and the solid was washed five times with water and dried overnight in vacuo. The obtained dark red powder was characterized as the mono-DMSO ligated complex **1DMSO**, **2DMSO**, **3DMSO** or **4DMSO**.

Synthesis of Bis-NH₂ Cobalt Corroles: The crude mixture, **Crude-Co**, obtained in the previous metalation step, was dissolved in 150 mL of dichloromethane, the organic phase washed four times (100 mL each) with an aqueous 0.4 M ammonia solution, followed by addition of five drops of 25 % aqueous ammonia to the solution. Methanol (50 mL) was added to the organic phase and dichloromethane removed, affording the desired product. The solid was then filtered and dried overnight. A dark green powder was obtained and characterized as the bis-NH₂ ligated cobalt corrole **1NH₂**, **2NH₂**, **3NH₂** or **4NH₂**.

(Cl₂Ph)₂CorCo(DMSO) (1DMSO): 89 % yield (124 mg). UV/Visible (toluene, 1 % DMSO): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 381 (51.9), 570 (10.6) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 9.13 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.62 (d, ³J_{H_{HH}} = 4.7 Hz, 2 H), 8.52 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.49 (d, ³J_{H_{HH}} = 4.7 Hz, 2 H), 7.74 (d, ³J_{H_{HH}} = 8.6 Hz, 6 H), 7.64–7.59 (m, 3 H), 2.60 (s, 6 H), –6.59 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 785.73 [M – DMSO]⁺, 785.89 calcd. for C₃₇H₁₇Cl₂CoN₄. HR-MS (ESI): *m/z* = 785.8933 [M – DMSO]⁺, 785.8911 calcd. for C₃₇H₁₇Cl₂CoN₄. DMSO content by TGA: Δm (calcd. for 1 DMSO ligated) = 11.46 % (10.10 %), T_{onset} = 124 °C, T_{end} = 225 °C.

(Cl₂Ph)₂CorCo(NH₂)₂ (1NH₂): 79 % yield (102 mg). UV/Visible (toluene, 1 % NH₄OH): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 440 (93.2), 451 (87.3), 514 (5.5), 589 (19.1), 617 (32.7) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 9.14 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.62 (d, ³J_{H_{HH}} = 4.6 Hz, 2 H), 8.53 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.49 (d, ³J_{H_{HH}} = 4.6 Hz, 2 H), 7.75 (m, 6 H), 7.62 (m, 3 H), –6.61 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 785.77 [M – 2NH₃]⁺, 785.89 calcd. for C₃₇H₁₇Cl₂CoN₄. MS (ESI) *m/z* = 820.30 [M]⁺, 819.94 calcd. for C₃₇H₂₃Cl₂CoN₆. HR-MS (ESI): *m/z* = 785.8943 [M – 2NH₃]⁺, 785.8911 calcd. for C₃₇H₁₇Cl₂CoN₄. NH₃ content by TGA: Δm (calcd. for 2 NH₃ ligated) = 4.66 % (4.16 %), T_{onset} = 113 °C, T_{end} = 176 °C. IR: $\bar{\nu}$ = 3360, 3260 cm⁻¹.

(CO₂MePh)Mes₂CorCo(DMSO) (2DMSO): 88 % yield (211 mg). UV/Visible (toluene, 1 % DMSO): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 383 (62.9), 563 (13.1) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 8.98 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.62 (d, ³J_{H_{HH}} = 4.7 Hz, 2 H), 8.59 (d, ³J_{H_{HH}} = 4.7 Hz, 2 H), 8.42 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.33 (d, ³J_{H_{HH}} = 7.9 Hz, 2 H), 8.26 (d, ³J_{H_{HH}} = 7.9 Hz, 2 H), 7.22 (s, 4 H), 4.02 (s, 3 H), 2.56 (s, 6

H), 2.55 (s, 6 H), 1.86 (s, 12 H), –6.68 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 723.75 [M – DMSO]⁺, 724.22 calcd. for C₄₅H₃₇CoN₄O₂. MS (ESI) *m/z* = 802.30 [M]⁺, 802.24 calcd. for C₄₇H₄₃CoN₄O₂S. HR-MS (ESI): *m/z* = 724.2263 [M – DMSO]⁺, 724.2243 calcd. for C₄₅H₃₇CoN₄O₂. DMSO content by TGA: Δm (calcd. for 1 DMSO ligated) = 12.29 % (10.78 %), T_{onset} = 144 °C, T_{end} = 208 °C.

(CO₂MePh)Mes₂CorCo(NH₂)₂ (2NH₂): 94 % yield (151 mg). UV/Visible (toluene, 1 % NH₄OH): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 439 (53.2), 461 (58.9), 514 (8.2), 584 (9.3), 623 (25.0) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 9.05 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.68 (d, ³J_{H_{HH}} = 4.6 Hz, 2 H), 8.66 (d, ³J_{H_{HH}} = 4.6 Hz, 2 H), 8.49 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.39 (d, ³J_{H_{HH}} = 8.2 Hz, 2 H), 8.32 (d, ³J_{H_{HH}} = 8.2 Hz, 2 H), 7.26 (s, 4 H), 4.07 (s, 3 H), 2.61 (s, 6 H), 1.90 (s, 12 H), –6.61 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 724.88 [M – 2NH₃]⁺, 724.22 calcd. for C₄₅H₃₇CoN₄O₂. MS (ESI) *m/z* = 758.26 [M]⁺, 758.28 calcd. for C₄₅H₄₃CoN₄O₂. HR-MS (ESI): *m/z* = 724.2244 [M – 2NH₃]⁺, 724.2243 calcd. for C₄₅H₃₇CoN₄O₂. NH₃ content by TGA: Δm (calcd. for 2 NH₃ ligated) = 5.41 % (4.76 %), T_{onset} = 100 °C, T_{end} = 175 °C. IR: $\bar{\nu}$ = 3338, 3257 cm⁻¹.

Ph₂CorCo(DMSO) (3DMSO): 71 % yield (70 mg). UV/Visible (toluene, 1 % DMSO): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 387 (77.0), 566 (14.1) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 9.07 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.91 (d, ³J_{H_{HH}} = 4.7 Hz, 2 H), 8.74–8.67 (m, 4 H), 8.22 (d, ³J_{H_{HH}} = 7.4 Hz, 4 H), 8.14 (d, ³J_{H_{HH}} = 7.4 Hz, 2 H), 7.72–7.56 (m, 9 H), 2.52 (s, 6 H), –6.47 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 581.71 [M – DMSO]⁺, 582.12 calcd. for C₃₇H₂₃CoN₄. MS (ESI) *m/z* = 660.33 [M]⁺, 660.14 calcd. for C₃₉H₂₉CoN₄O₂S. HR-MS (ESI): *m/z* = 582.1257 [M – DMSO]⁺, 582.1249 calcd. for C₃₇H₂₃CoN₄. DMSO content by TGA: Δm (calcd. for 1 DMSO ligated) = 10.53 % (11.20 %), T_{onset} = 120 °C, T_{end} = 193 °C.

Ph₂CorCo(NH₂)₂ (3NH₂): 80 % yield (237 mg). UV/Vis (toluene, 1 % NH₄OH): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 436 (29.0), 455 (25.0), 526 (3.0), 585 (4.2), 629 (12.0) nm. ¹H NMR [500 MHz, CDCl₃ + NH₃ (g)]: δ = 9.08 (d, ³J_{H_{HH}} = 3.2 Hz, 2 H), 8.91 (d, ³J_{H_{HH}} = 4.2 Hz, 2 H), 8.72 (m, 4 H), 8.23 (d, ³J_{H_{HH}} = 7.4 Hz, 4 H), 8.15 (d, ³J_{H_{HH}} = 7.4 Hz, 2 H), 7.64 (m, 9 H), –6.46 (s, 6 H) ppm. MS (MALDI-TOF) *m/z* = 582.64 [M – 2NH₃]⁺, 582.13 calcd. for C₃₇H₂₃CoN₄. MS (ESI) *m/z* = 616.20 [M]⁺, 616.18 calcd. for C₃₇H₂₉CoN₄O₂S. HR-MS (ESI): *m/z* = 582.1252 [M – 2NH₃]⁺, 582.1249 calcd. for C₃₇H₂₃CoN₄. NH₃ content by TGA: Δm (calcd. for 2 NH₃ ligated) = 4.45 % (4.72 %), T_{onset} = 100 °C, T_{end} = 172 °C. IR: $\bar{\nu}$ = 3338 cm⁻¹, 3257 cm⁻¹. IR: 3413, 3252 cm⁻¹.

[(OMe)₂Ph]Mes₂CorCo(DMSO) (4DMSO): 83 % yield (193 mg). UV/Visible (toluene, 1 % DMSO): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 380 (63.4), 564 (14.1) nm. ¹H NMR [500 MHz, CDCl₃ + 1 % NH₂NH₂·H₂O]: δ = 9.00 (d, ³J_{H_{HH}} = 4.0 Hz, 2 H), 8.55 (d, ³J_{H_{HH}} = 4.5 Hz, 2 H), 8.49 (d, ³J_{H_{HH}} = 4.5 Hz, 2 H), 8.44 (d, ³J_{H_{HH}} = 4.0 Hz, 2 H), 7.25 (s, 4 H), 6.54 (s, 2 H), 4.07 (s, 3 H), 3.48 (s, 6 H), 2.60 (m, 12 H), 1.88 (s, 12 H), –3.89 (s, 4 H) ppm. MS (MALDI-TOF) *m/z* = 756.09 [M – DMSO]⁺, 756.25 calcd. for C₄₆H₄₁CoN₄O₃. MS (ESI) *m/z* = 834.34 [M]⁺, 834.27 calcd. for C₄₈H₄₇CoN₄O₄S. HR-MS (ESI): *m/z* = 756.2507 [M – DMSO]⁺, 756.2505 calcd. for C₄₆H₄₁CoN₄O₃. DMSO content by TGA: Δm (calcd. for 1 DMSO ligated) = 9.88 % (9.60 %), T_{onset} = 158 °C, T_{end} = 233 °C.

[(OMe)₂Ph]Mes₂CorCo(NH₂)₂ (4NH₂): 84 % yield (180 mg). UV/Visible (toluene, 1 % NH₄OH): λ_{max} ($\epsilon \times 10^{-3}$ L·mol⁻¹ cm⁻¹) 434 (55.3), 453 (55.9), 514 (6.2), 585 (10.0), 624 (24.9) nm. ¹H NMR [500 MHz, CDCl₃ + 1 % NH₂NH₂·H₂O]: δ = 9.05 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 8.61 (d, ³J_{H_{HH}} = 4.5 Hz, 2 H), 8.55 (d, ³J_{H_{HH}} = 4.5 Hz, 2 H), 8.49 (d, ³J_{H_{HH}} = 4.1 Hz, 2 H), 7.29 (s, 4 H), 6.58 (s, 2 H), 4.11 (s, 3 H), 3.52 (s, 6 H), 2.64 (s, 6 H), 1.93 (s, 12 H), –4.18 (s, 4 H) ppm. MS (MALDI-TOF) *m/z* = 756.08 [M – 2NH₃]⁺, 756.25 calcd. for C₄₆H₄₁CoN₄O₃. MS (ESI) *m/z* =

790.30 [M]⁺, 790.30 calcd. for C₄₆H₄₇CoN₆O₅. HR-MS (ESI): *m/z* = 756.2507 [M – 2NH₃]⁺, 756.2505 calcd. for C₄₆H₄₁CoN₆O₅. NH₃ content by TGA: Δ*m* (calcd. for 2 NH₃ ligated) = 3.12 % (3.60 %), *T*_{onset} = 42 °C, *T*_{end} = 120 °C. IR: ν = 3357, 3261 cm^{−1}.

Crystal Structure Determination of 4H and 1NH₂. Light violet needle-shaped crystals of free-base **4H** were recrystallized from a mixture of water and acetone by slow evaporation and a suitable crystal (0.25 × 0.20 × 0.08) mm³ was mounted with MITIGEN holder oil on a Bruker D8 VENTURE diffractometer and kept at 100(1) K during data collection. For the cobalt corrole **1NH₂**, clear dark violet plate-shaped crystals were obtained from a mixture of hexane and DCM by slow evaporation and a suitable crystal (0.22 × 0.20 × 0.15) mm³ was selected for structure determination. Both crystals were examined with Olex²¹⁴³ and the structures solved with the ShelXT¹⁷⁵³ structure solution program, using the Intrinsic Phasing solution method. The model was refined with version 2016/6 of ShelXL¹⁷⁴³ using least-squares minimization. Additional details are given in supporting information.

Crystal Data for Free-base Corrole 4H: C₄₉H₅₀N₄O₄, *M_r* = 758.93, monoclinic, *P*2₁/c (No. 14), *a* = 7.5501(2) Å, *b* = 24.8197(8) Å, *c* = 21.8942(7) Å, β = 97.165(2)°, α = γ = 90°, *V* = 4070.7(2) Å³, *T* = 100(1) K, *Z* = 4, *Z'* = 1, μ(Cu–Kα) = 0.624, 35965 reflections measured, 7175 unique (*R*_{int} = 0.0982) which were used in all calculations. The final *wR*₂ was 0.1167 (all data) and *R*₁ was 0.0556 [*I* > 2(*I*)].

Crystal Data for 1NH₂: C₇₇H₅₃Cl₁₂CoN₁₂, *M_r* = 1689.57, monoclinic, *P*2₁/c (No. 14), *a* = 12.7201(5) Å, *b* = 7.3022(3) Å, *c* = 38.3166(16) Å, β = 94.677(2)°, α = γ = 90°, *V* = 3547.2(3) Å³, *T* = 100(1) K, *Z* = 2, *Z'* = 0.5, μ(Cu–Kα) = 8.272, 21938 reflections measured, 6253 unique (*R*_{int} = 0.0825) which were used in all calculations. The final *wR*₂ was 0.2447 (all data) and *R*₁ was 0.0955 [*I* > 2(*I*)].

Supporting Information (see footnote on the first page of this article): UV/Visible spectra, cyclic voltammograms, MALDI/TOF MS spectra, ¹H NMR, ESR spectra and CIF file giving crystallographic data for compound **1NH₂** and compound free-base corrole **4H**, selected bond lengths and angles as well as XYZ coordinates (DFT calculations).

CCDC 1812155 (for **1NH₂**), and 181709 (for free-base corrole **4H**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre.

Notes

The authors declare no competing financial interest.

Conflict of Interest

The authors declare no conflict of interest.

Acknowledgments

This work was supported by the Robert A. Welch Foundation (K. M. K., Grant E-680). Support was provided by the Centre National de la Recherche Scientifique (CNRS) (UMR UB-Centre National de la Recherche Scientifique (CNRS) 6302), the "Université de Bourgogne Franche-Comté", the FEDER-FSE Bourgogne 2014/2020 (European Regional Development Fund) and the "Conseil Régional de Bourgogne" through the PARI II CDEA project and the JCE program. We are thankful to Mrs. Sandrine

Pacquelet for technical assistance and W. Ryan Osterloh for experimental help. ANR is granted for financial support (CO3SENS). The authors wish also to warmly thank Prof. Anthony Romieu (IUF, UBFC, ICMUB, UMR CNRS 6302) for recording the LRMS ESI spectra of cobalt corrole complexes.

Keywords: Cobalt · DMSO binding · Ammonia binding · Electrochemistry · Coordination chemistry

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Received: July 18, 2018