

# Assessing the Performances of CASPT2 and NEVPT2 for Vertical Excitation Energies

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## Abstract

Amongst the available theories for accurately modeling electronic excited states, methods able to simultaneously account for both static and dynamic electron correlations have been considered, not only to model photochemical events, but also to provide reference values for vertical transition energies, hence allowing to benchmark lower-order models. In this category, both the complete-active-space second-order perturbation theory (CASPT2) and the  $N$ -electron valence state second-order perturbation theory (NEVPT2) are certainly popular, the latter presenting the advantage of not requiring the application of the empirical ionization-potential-electron-affinity (IPEA) and level shifts. However, the actual accuracy of these multiconfigurational approaches is not settled yet. In this context, the present work relies on highly-accurate *aug-cc-pVTZ* vertical transition energies for 284 excited states of diverse character (174 singlet, 110 triplet, 206 valence, 78 Rydberg, 78  $n \rightarrow \pi^*$ , 119  $\pi \rightarrow \pi^*$ , and 8 double excitations) determined in 35 small- to medium-sized organic molecules containing from three to six non-hydrogen atoms to assess the performances of these approaches. The CASPT2 calculations are performed with and without IPEA shift and compared to the partially-contracted (PC) and strongly-contracted (SC) variants of NEVPT2. We find that both CASPT2 with IPEA shift and PC-NEVPT2 provide fairly reliable vertical transition energy estimates, with slight overestimations and mean absolute errors of 0.11 and 0.13 eV, respectively. These values are found to be rather uniform for the various subgroups of transitions. The present work completes our previous benchmarks focussed on single-reference wave function methods (*J. Chem. Theory Comput.* **14**, 4360 (2018); *ibid.*, **16**, 1711 (2020)), hence allowing for a fair comparison between various families of electronic structure methods. In particular, we show that ADC(2), CCSD, and CASPT2 deliver similar accuracies for excited states with a dominant single-excitation character.

## 1. INTRODUCTION

In the early 90's, it was certainly challenging to systematically introduce dynamical electron correlation effects in the description of molecular systems, so as to access accurate geometries, dissociation energies, vibrational frequencies, spectroscopic constants, as well as vertical transition energies (VTEs). In this framework, the development and efficient implementation of single-reference many-body perturbation theory (SR-MBPT) based on a Hartree-Fock (HF) reference wave function were certainly major steps forward for the electronic structure community. At that time, most applications using SR-MBPT were based on the acclaimed second-order Møller-Plesset perturbative correction,<sup>1</sup> and while the reliability of such approach was found to be quite satisfactory for closed-shell systems, it obviously suffers from major drawbacks when the HF reference wave function is not a valid starting point, i.e., when the targeted electronic state is not properly described by a single Slater determinant. This situation is ubiquitous, for example, in transition metal complexes, bond-breaking reactions, transition states, and, more importantly in the present context, electronic excited states

(ESs). This severe limitation has led to the development of multireference extensions of MBPT (MR-MBPT) in which the reference wave function is multiconfigurational, i.e., contains more than one Slater determinant. These powerful approaches allow to account for both static electron correlation encapsulated in the reference multiconfigurational wave function — most often described within the complete-active-space self-consistent field (CASSCF) formalism — and dynamic electron correlation treated very efficiently at low order of perturbation theory.

Various flavors and implementations of MR-MBPT have emerged over the years. Among the most well-known variants, one can cite the complete-active-space second-order perturbation theory (CASPT2) developed by Roos and coworkers,<sup>2,3</sup> the multireference second-order Møller-Plesset (MRMP2) approach proposed by Hirao,<sup>4</sup> and the  $N$ -electron valence state second-order perturbation theory (NEVPT2) designed by Angeli, Malrieu, and coworkers.<sup>5-7</sup> All these methods have proven to be efficient and accurate when one deals with ESs, in particular for the computation of VTEs. Illustratively, three decades after its original formulation,

CASPT2 has clearly become the most popular multiconfigurational approach of molecular quantum chemistry in order to tackle ground states and ESs of multiconfigurational character on an equal footing.

Nevertheless, it is noteworthy that several limitations of CASPT2 were rapidly disclosed and circumvented. The first recognized drawback of CASPT2 in the context of ESs was the appearance of intruder states, i.e., perturbers that have nearly equal energies to the zeroth-order CASSCF wave function.<sup>8–10</sup> The best remedy to this severe issue is to move the intruder states in the active space but this strategy is obviously intrinsically limited by the dreadful computational cost associated with the expansion of the active space. An efficient alternative was formulated by Roos and Andersson<sup>11</sup> who proposed to introduce a real-valued shift in the energy denominators of the second-order energy (thus avoiding singularities) and correcting accordingly the resulting energy. This approach, known as *level-shift correction*, is almost systematically applied nowadays. Note that, in the case of “weak” intruder states, an imaginary shift can be also introduced.<sup>12</sup>

The second bottleneck of the single-state (also called state-specific) CASPT2 (SS-CASPT2) method, as implemented originally by Roos and coworkers, is the difficulty to deal with electronic state mixing (e.g., valence/Rydberg and covalent/ionic mixing). A solution consists in allowing the states to mix within the MR-MBPT treatment, giving rise to the multistate formulation of CASPT2 denoted as MS-CASPT2,<sup>13</sup> and later refined in an extended multistate version called XMS-CASPT2.<sup>14</sup> Very recently, a further extension known as extended dynamically weighted CASPT2 (XDW-CASPT2) was proposed in order to combine the most attractive features of both MS-CASPT2 and XMS-CASPT2.<sup>15</sup>

The third bottleneck was found in evaluating a large number of chemical problems for which systematic errors were noticed<sup>16,17</sup> and ascribed to the unbalanced description of the zeroth-order Hamiltonian for the open- and closed-shell electronic configurations. This systematic error can be attenuated by introducing an additional parameter, the so-called ionization-potential-electron-affinity (IPEA) shift, in the zeroth-order Hamiltonian.<sup>18</sup>

Whilst these three “fixes” improve the overall accuracy of CASPT2, all of them remain rather empirical, which can be viewed as a drawback compared to other *ab initio* approaches. In this context, NEVPT2 has the advantage of being practically free of intruder states and enjoys the valuable property of being size-consistent (which is not the case of CASPT2). NEVPT2 mainly differs from CASPT2 in the choice of (i) the zeroth-order Hamiltonian, and (ii) the definition of the perturbers and their corresponding energies. NEVPT2 exists in two different contraction schemes: strongly-contracted (SC) and partially-contracted (PC), the latter being supposedly more accurate thanks to its larger number of perturbers and its greater flexibility. Note that a multistate version of NEVPT2 known as quasi-degenerate NEVPT2 (QD-NEVPT2) has also been developed.<sup>19</sup>

Both the parent and stepwisely improved versions of CASPT2 have been applied to tackle a wide variety of chemical problems: heavy element chemistry,<sup>20</sup> biochem-

ical systems,<sup>21,22</sup> transition metal complexes<sup>23–29</sup> (such as bimetallic complexes<sup>30</sup>), and, of course, ESs. The latter were intensively investigated by Roos, Serrano-Andrés, and their collaborators from the start,<sup>31–43</sup> in works where experiment was typically used to appraise the quality of the CASPT2 estimates. These estimates were themselves used later as references in benchmark studies assessing the performances of “lower-level” ES methods.<sup>44–47</sup> Starting in 2008, very comprehensive benchmarks of valence VTEs in small and medium CNOH compounds were performed by Thiel’s group.<sup>48–53</sup> Besides literature data, these authors first relied on CASPT2/TZVP to define a list of 104 theoretically best estimates (TBEs) for the singlet vertical transitions in 28 organic molecules optimized at the MP2/6-31G(d) level. In their original study,<sup>48</sup> the TBEs associated with 63 triplet ESs were obtained at the third-order coupled-cluster (CC3)<sup>54</sup> level with the same TZVP basis set. These different choices of reference methods for the two ES families were justified by the almost pure single-excitation character of the triplet transitions, whereas the singlet transitions showed a less clear-cut nature. In 2010, Thiel’s group upgraded their TBEs using the *aug-cc-pVTZ* basis set<sup>52,53</sup> and found that using this larger basis set containing additional diffuse functions downshifted the singlet VTEs by an average of 0.11 eV as compared to TZVP, with a high degree of consistency between the two sets of data (correlation coefficient larger than 0.996).<sup>52</sup> Interestingly, in their studies employing the *aug-cc-pVTZ* basis set, Thiel and coworkers also shifted to CC3 as the default reference approach for defining the TBEs associated with singlet transitions. Yet, no error bar was defined, nor estimated for CASPT2.

Both the original TZVP and the improved *aug-cc-pVTZ* TBEs of Thiel and coworkers were later employed in countless benchmarks of ES models.<sup>49–51,55–78</sup> In particular, we wish to pinpoint valuable studies of NEVPT2/CASPT2<sup>68</sup> and CASSCF<sup>78</sup> using Thiel’s TZVP reference values. These works are clearly in the same philosophy as the present effort, though these were obviously limited to valence transitions of single-excitation nature. At the same time, the number of ES studies employing CASPT2 steadily increased during the last two decades, with investigations of many photophysical and photochemical processes: photosensitization, photoisomerization, charge transfer, nonradiative deactivation, as well as photodynamics studies of small to medium size organic molecules.<sup>79–108</sup>

Given such extensive use of CASPT2 by both the electronic structure and dynamics communities, González and co-workers proposed an excellent reassessment of this approach in 2017.<sup>109</sup> In this key work, they showed that, for di- and tri-atomic molecules for which full configuration interaction (FCI) energies could be computed (albeit with a small basis set), standard CASPT2<sup>3</sup> slightly underestimates the VTEs, while the application of the IPEA shift<sup>18</sup> partially corrects this underestimation. The same work also reports that applying an IPEA shift can lead to overestimations of the VTEs for medium-sized molecules, and even quite significant exaggerations for larger molecules (of the same order of magnitude as the underestimation obtained without IPEA for the smallest derivatives). As a result, the authors of Ref.

109 concluded that the application of the IPEA shift is not systematically justified for organic chromophores. However, and importantly, this work also shows that the relevance of applying IPEA is basis set dependent: for double- $\zeta$  basis sets, smaller errors are obtained when setting the IPEA shift to zero, whereas for more extended basis set, of triple- or quadruple- $\zeta$  quality, applying an IPEA shift improves the accuracy. In such context, one can find several other examples in the literature showing the importance of introducing an IPEA shift to reach a better agreement between experiment and theory, as for, e.g., the excitation energies in iron complexes,<sup>24,27,29,110–113</sup> or in BODIPY derivatives.<sup>114</sup>

Of course, besides CASPT2 and related MR-MBPT, there exist many alternative *ab initio* methods for tackling ESs. In this regard, one should certainly cite several efficient single-reference approaches for modeling VTEs in large systems: (i) time-dependent density-functional theory (TD-DFT),<sup>115,116</sup> (ii) the Bethe-Salpeter equation (BSE) formalism,<sup>117,118</sup> (iii) the second-order algebraic-diagrammatic construction [ADC(2)],<sup>119</sup> and (iv) the second-order coupled-cluster (CC2) model.<sup>120,121</sup> However, the overall accuracy provided by these four methodologies is typically insufficient to be considered as a safe reference for assessing other methods.<sup>122</sup> We underline here that in contrast to TD-DFT and BSE, both the ADC and CC advantageously offer a path for systematic improvement through the increase of the expansion order, with, e.g., third-order approaches like ADC(3)<sup>70</sup> and CC3.<sup>54</sup> In this context, some of us, inspired by Thiel’s works, have recently proposed a new set of highly-accurate TBEs for VTEs of a large variety of ESs in small- and medium-sized organic molecules and radicals.<sup>123–130</sup> These TBEs were obtained directly from FCI using a selected CI approach,<sup>131,132</sup> CCSDTQ,<sup>133–136</sup> and CCSDT,<sup>137–141</sup> for compounds containing 1-to-3, 4, and 5-to-10 non-hydrogen atoms, respectively. Taking advantages of these highly trustworthy values, we were able to resolve some unanswered questions about the relative accuracies of ADC(3),<sup>70</sup> CCSDT-3,<sup>142,143</sup> and CC3.<sup>54</sup> In particular, we evidenced that the accuracy of the VTEs obtained with CCSDTQ is comparable to that of FCI, whereas both CC3 and CCSDT-3 can be viewed as highly reliable, at least for transitions with a predominant single-excitation character.<sup>123–129</sup>

In the present study, considering the ESs studied in Refs. 123 and 125 which constitutes a set of 284 ESs (174 singlet, 110 triplet, 206 valence, 78 Rydberg, 78  $n \rightarrow \pi^*$ , 119  $\pi \rightarrow \pi^*$ , and 8 double excitations) in 35 organic molecules containing from three to six non-hydrogen atoms (see Figure 1), we perform a comprehensive benchmark of both CASPT2 and NEVPT2, the former being considered with and without IPEA shift. Because we rely on very high-quality reference values, we believe that the present study can provide definite answers to the question of the relative accuracy of these multireference approaches for VTEs dominated by a single-excitation character, as well as reliable and trustworthy comparisons with single-reference ES methods.

## 2. COMPUTATIONAL DETAILS

We have computed VTEs for ESs of each compound displayed in Figure 1 using both CASPT2<sup>2,3,11,18</sup> and

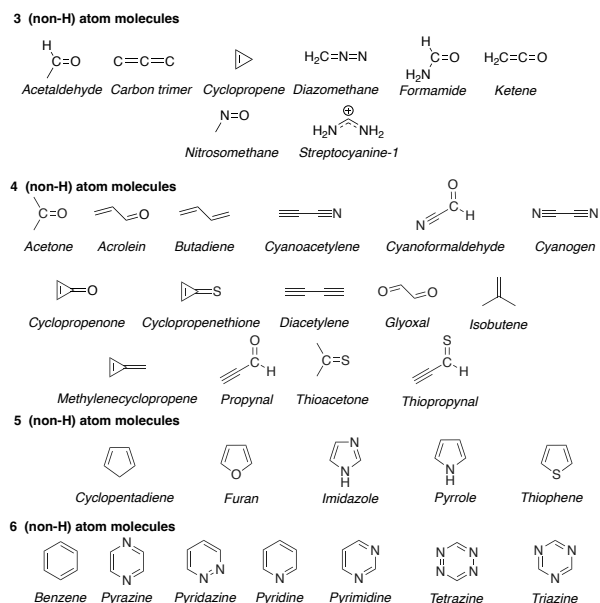


Figure 1: Various families of compounds considered in this study.

NEVPT2,<sup>5–7</sup> with Dunning’s *aug-cc-pVTZ* basis set.<sup>144</sup> Note that while all CASPT2 calculations are new, several (but not all) PC-NEVPT2 data were taken from a previous work.<sup>125</sup> The cartesian coordinates of the ground-state geometries of all considered molecules are taken from our previous studies.<sup>123–125</sup> These high-quality ground-state equilibrium geometries have been computed at the CC3<sup>54</sup> level, several of them being extracted from earlier works.<sup>145–147</sup>

All perturbative calculations reported herein were performed on top of state-averaged (SA) CASSCF wave functions, which were produced by considering the ground state and (at least) the ES of interest. Sometimes, more ESs were included in the state-averaging procedure to guarantee convergence and avoid root-flipping problems. The relevant information concerning the construction of the SA-CASSCF wave functions can be found in the Supporting Information (SI) of the present article for all electronic states not reported in Refs. 124 and 125, whereas the corresponding information for all other electronic states can be found in the SI of these references.

We tackled the intruder state problem both by increasing the size of active spaces as well as by introducing a level shift (0.3 a.u. unless otherwise stated) as discussed in Ref. 11. One set of CASPT2 calculations was performed by introducing the standard IPEA parameter (0.25 a.u.) [defined below as CASPT2(IPEA)] as discussed in Ref. 18 and another set of calculations was performed without IPEA correction [defined below as CASPT2(NOIPEA)]. We have performed both SC-NEVPT2 and PC-NEVPT2 calculations, and report only the latter in the main text below. The interested reader can find SC-NEVPT2 results in the SI. All the SA-CASSCF, CASPT2 (RS2 contraction level)<sup>148</sup> and NEVPT2 calculations were performed using the MOLPRO program.<sup>149</sup>

In the statistical analysis presented below, we report the usual statistical indicators: the mean signed error (MSE),



the mean absolute error (MAE), the root-mean-square error (RMSE), the standard deviation of the errors (SDE), as well as largest positive and negative deviations [Max(+) and Max(-), respectively].

### 3. RESULT AND DISCUSSION

Below, we first discuss individual molecules, trying to pinpoint relevant examples, before moving to statistical analyses. We note here that, as general trends will emerge during the discussion, we have particularly detailed the discussion on the first compounds and tried not to repeat general statements for each of the molecules regarding, e.g., relative accuracies for Rydberg and valence transitions, or the impact of IPEA. The full list of results can be found in Tables 1, 2 and 3 whereas SA-CASSCF and SC-NEVPT2 results can be found in Section S3 of the SI. Except when specifically discussed below, all the ESs considered in these Tables present a strongly dominant single-excitation character.

#### 3.1 Molecules with three non-hydrogen atoms

We have estimated VTEs in eight molecules containing three non-H atoms, namely, acetaldehyde, carbon trimer, cyclopropene, diazomethane, formamide, ketene, nitrosomethane, and streptocyanine-C1. Reference FCI (or high-level CC) values with various basis sets, as well as literature data can be found in our previous works.<sup>123,124</sup> The choice of active spaces and state-averaging procedure are detailed in Tables S1–S7 in the SI, except for carbon trimer where it was described elsewhere.<sup>124</sup>

*Acetaldehyde.* The VTEs computed for the lowest  $n \rightarrow \pi^*$  singlet and triplet transitions of acetaldehyde obtained with NEVPT2 at 4.39 and 4.00 eV, respectively, slightly exceed older estimates obtained at the same level of theory (4.29 and 3.97 eV),<sup>150</sup> and both are fitting the TBEs of 4.31 and 3.97 eV, obtained on the basis of FCI calculations.<sup>123</sup> Likewise CASPT2(IPEA) estimates are also accurate at 4.35 and 3.94 eV, whereas the absence of IPEA shift produces underestimated values (4.13 and 3.71 eV).

*Carbon trimer.* This original linear molecule is of interest because it presents two “pure” doubly-excited states for which FCI-quality TBEs are available. As one can be seen from the data listed of Table 1, NEVPT2 provides values extremely close to FCI, whereas the absence of IPEA shift yields very large underestimations.

*Cyclopropene.* The lowest two singlet  $\sigma \rightarrow \pi^*$  and  $\pi \rightarrow \pi^*$  transitions of cyclopropene, of respective  $B_1$  and  $B_2$  symmetries, are very close in energy,<sup>48,123</sup> an effect reproduced by all second-order perturbation theories considered here. While both CASPT2(IPEA) and NEVPT2 deliver the same ordering as CCSDT,<sup>123</sup> CASPT2(NOIPEA) swaps the two states and it significantly underestimates the VTE associated with the  $\pi \rightarrow \pi^*$  excitation. All VTEs obtained with CASPT2(IPEA) are within 0.1 eV of the TBEs, except for the singlet  $\sigma \rightarrow \pi^*$  transition for which the error is slightly larger (i.e., within 0.2 eV). Switching off the IPEA shift downshifts all CASPT2 estimates from 0.2 eV up to 0.4 eV, leading to larger errors.

*Diazomethane.* The  $\pi \rightarrow \pi^*$  transition energies of both singlet and triplet symmetries of diazomethane obtained from CASPT2(IPEA) are very close to the TBEs, except for the

valence  $^1A_1(\pi \rightarrow \pi^*)$  transition for which an error of 0.3 eV is obtained. A similar observation can be made for NEVPT2. VTEs for Rydberg excitations are also nicely reproduced by both methods with errors around 0.1 eV except for the  $^3A_1(R; \pi \rightarrow 3p)$  transition for which the error increases to 0.2 eV. CASPT2(NOIPEA) underestimates all the TBEs, apart from the valence  $^1A_1(\pi \rightarrow \pi^*)$  transition which is overestimated upon inclusion of the IPEA correction. For the eight transitions considered the MAEs are 0.09, 0.15, and 0.11 eV for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively.

*Formamide.* The lowest singlet and triplet  $n \rightarrow \pi^*$  transitions of formamide obtained with CASPT2(IPEA) are within 0.02 eV of our TBEs, whereas the CASPT2(NOIPEA) values are too small. NEVPT2 is also very accurate for these two transitions. For the three higher singlet states of  $A'$  symmetry, the presence of a strong state mixing between the valence and Rydberg transitions was observed in previous CC studies by us<sup>123</sup> and the Szalay’s group.<sup>151</sup> This mixing led to a problematic assignment of these states, as the lowest Rydberg state had a larger oscillator strength than the valence ( $\pi \rightarrow \pi^*$ ) state, which is counterintuitive. Our SA-CASSCF reference wave function shows that such a mixing is weak and the assignment of the states is straightforward, highlighting a clear advantage of multi-reference methods in this case. The oscillator strengths are also consistent with the electronic nature of the states, the valence ( $\pi \rightarrow \pi^*$ ) state having the largest transition dipole moment (3.2 D) compared to the Rydberg states (0.9 and 1.6 D). Overall, the results obtained with CASPT2(IPEA) are consistent with previous TBEs,<sup>123</sup> although we do not consider these transitions in the statistics below due to the state-mixing issue.

*Ketene.* For the eight ESs of ketene listed in Table 2, the MAEs are 0.05, 0.16, and 0.07 eV for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively. As in the previous systems, CASPT2(NOIPEA) underestimates systematically all the TBEs. Ketene was previously investigated at the MS-CASPT2/6-31+G(d) level by the Morokuma group,<sup>152</sup> who reported VTEs of 3.72, 5.97, 3.62, 5.42, and 5.69 eV for the  $^1A_2$ ,  $^1B_1$ ,  $^3A_2$ ,  $^3A_1$ , and  $^3B_1$  excitations, respectively, all values being slightly below the TBEs. Despite the different basis sets used, one also notes a reasonable match between these previous MS-CASPT2 values and the present results that rely on rather comparable active spaces.

*Nitrosomethane.* Both the lowest-lying singlet and triplet transitions of nitrosomethane are of  $n \rightarrow \pi^*$  character and the NEVPT2 VTEs are in very good agreement with the TBEs, whereas CASPT2 provides too small values. More interestingly, the second singlet ES has a pure double electronic excitation of  $n, n \rightarrow \pi^*, \pi^*$  nature, and a FCI value of 4.76 eV could be obtained.<sup>124</sup> In contrast to single-reference methods that miserably fail for this transition except when quadruples are included,<sup>124</sup> all three second-order multireference perturbative methods tested here deliver very accurate results with a slight underestimation by less than 0.1 eV.

*Streptocyanine-C1.* The shortest cyanine is an interesting case as its transition cannot be accurately described by TD-DFT.<sup>153,154</sup> At the CASPT2 level, our values are comparable to the ones obtained by Send *et al.*,<sup>153</sup> both with and with-

Table 1: Comparisons between TBEs taken from Table 6 of Ref. 123, Table 1 of Ref. 124, and Table 11 of Ref. 125 and VTEs computed at the CASPT2(IPEA), CASPT2(NOIPEA), and (PC-)NEVPT2 levels. All values are given in eV and computed with the *aug-cc-pVTZ* basis.

Compound	State	CASPT2			Compound	State	CASPT2			
		TBE	IPEA	NOIPEA			TBE	IPEA	NOIPEA	
Acetaldehyde	$^1A''(V; n \rightarrow \pi^*)$	4.31	4.35	4.13	Cyclopropenone	$^1B_1(V; n \rightarrow \pi^*)$	4.26	4.12	3.75	4.04
	$^3A''(V; n \rightarrow \pi^*)$	3.97	3.94	3.71		$^1A_2(V; n \rightarrow \pi^*)$	5.55	5.62	5.31	5.85
Acetone	$^1A_2(V; n \rightarrow \pi^*)$	4.47	4.44	4.19	$^1B_2(R; n \rightarrow 3s)$	6.34	6.28	6.21	6.51	
	$^1B_2(R; n \rightarrow 3s)$	6.46	6.46	6.35	$^1B_2(V; \pi \rightarrow \pi^*)$	6.54	6.84	6.70	6.82	
	$^1A_2(R; n \rightarrow 3p)$	7.47	7.80	7.55	$^1B_2(R; n \rightarrow 3p)$	6.98	6.54	6.20	7.07	
	$^1A_1(R; n \rightarrow 3p)$	7.51	7.67	7.46	$^1A_1(R; n \rightarrow 3p)$	7.02	7.27	7.03	7.28	
	$^1B_2(R; n \rightarrow 3p)$	7.62	7.56	7.47	$^1A_1(V; \pi \rightarrow \pi^*)$	8.28	8.90	8.71	8.19	
	$^3A_2(V; n \rightarrow \pi^*)$	4.13	4.13	3.89	$^3B_1(V; n \rightarrow \pi^*)$	3.93	3.65	3.28	3.51	
	$^3A_1(V; \pi \rightarrow \pi^*)$	6.25	6.24	6.07	$^3B_2(V; \pi \rightarrow \pi^*)$	4.88	4.76	4.60	5.10	
Acrolein	$^1A''(V; n \rightarrow \pi^*)$	3.78	3.70	3.48	$^3A_2(V; n \rightarrow \pi^*)$	5.35	5.36	5.06	5.60	
	$^1A'(V; \pi \rightarrow \pi^*)$	6.69	6.93	6.37	$^3A_1(V; \pi \rightarrow \pi^*)$	6.79	6.93	6.61	7.16	
	$^1A''(V; n \rightarrow \pi^*)$	6.72	7.12	5.97	Cyclopropenethione	$^1A_2(V; n \rightarrow \pi^*)$	3.41	3.43	3.14	3.52
	$^1A'(R; n \rightarrow 3s)$	7.08	7.12	6.90		$^1B_1(V; n \rightarrow \pi^*)$	3.45	3.45	3.17	3.50
	$^1A'(V; \pi \rightarrow \pi^*)^a$	7.87	8.18	7.84	$^1B_2(V; \pi \rightarrow \pi^*)$	4.60	4.64	4.35	4.77	
	$^3A''(V; n \rightarrow \pi^*)$	3.51	3.42	3.19	$^1B_2(R; n \rightarrow 3s)$	5.34	5.25	5.15	5.35	
	$^3A'(V; \pi \rightarrow \pi^*)$	3.94	4.03	3.78	$^1A_1(V; \pi \rightarrow \pi^*)$	5.46	5.84	5.32	5.54	
	$^3A'(V; \pi \rightarrow \pi^*)$	6.18	6.29	5.93	$^1B_2(R; n \rightarrow 3p)$	5.92	5.93	5.86	5.99	
	$^3A''(V; n \rightarrow \pi^*)$	6.54	6.81	5.47	$^3A_2(V; n \rightarrow \pi^*)$	3.28	3.28	3.00	3.38	
	Benzene	$^1B_{2u}(V; \pi \rightarrow \pi^*)$	5.06	5.14	4.66	$^3B_1(V; n \rightarrow \pi^*)$	3.32	3.35	3.07	3.40
$^1B_{1u}(V; \pi \rightarrow \pi^*)$		6.45	6.65	6.23	$^3B_2(V; \pi \rightarrow \pi^*)$	4.01	3.97	3.75	4.17	
$^1E_{1g}(R; \pi \rightarrow 3s)$		6.52	6.70	6.57	$^3A_1(V; \pi \rightarrow \pi^*)$	4.01	4.01	3.77	4.13	
$^1A_{2u}(R; \pi \rightarrow 3p)$		7.08	7.21	7.06	Diacetylene	$^1\Sigma_u^-(V; \pi \rightarrow \pi^*)$	5.33	5.42	5.01	5.33
$^1E_{2u}(R; \pi \rightarrow 3p)$		7.15	7.26	7.11		$^1\Delta_u(V; \pi \rightarrow \pi^*)$	5.61	5.68	5.30	5.61
$^1E_{2g}(V; \pi \rightarrow \pi^*)^a$		8.28	8.31	7.82 <sup>b</sup>	$^3\Sigma_u^+(V; \pi \rightarrow \pi^*)$	4.10	4.11	3.67	4.08	
$^3B_{1u}(V; \pi \rightarrow \pi^*)$		4.16	4.22	3.92	$^3\Delta_u(V; \pi \rightarrow \pi^*)$	4.78	4.82	4.45	4.78	
$^3E_{1u}(V; \pi \rightarrow \pi^*)$		4.85	4.89	4.51	Diazomethane	$^1A_2(V; \pi \rightarrow \pi^*)$	3.14	3.13	2.92	3.09
$^3B_{2u}(V; \pi \rightarrow \pi^*)$		5.81	5.85	5.40		$^1B_1(R; \pi \rightarrow 3s)$	5.54	5.50	5.30	5.63
Butadiene		$^1B_u(V; \pi \rightarrow \pi^*)$	6.22	6.76	6.52	$^1A_1(V; \pi \rightarrow \pi^*)$	5.90	6.21	5.92	6.23
	$^1B_g(R; \pi \rightarrow 3s)$	6.33	6.49	6.31	$^3A_2(V; \pi \rightarrow \pi^*)$	2.79	2.87	2.67	2.83	
	$^1A_g(V; \pi \rightarrow \pi^*)^a$	6.50	6.74	6.29	$^3A_1(V; \pi \rightarrow \pi^*)$	4.05	4.10	3.88	4.07	
	$^1A_u(R; \pi \rightarrow 3p)$	6.64	6.74	6.64	$^3B_1(R; \pi \rightarrow 3s)$	5.35	5.34	5.15	5.48	
	$^1A_u(R; \pi \rightarrow 3p)$	6.80	6.95	6.84	$^3A_1(R; \pi \rightarrow 3p)$	6.82	7.00	6.76	7.01	
	$^1B_u(R; \pi \rightarrow 3p)$	7.68	7.60	7.30	$^1A''[F](V; \pi \rightarrow \pi^*)$	0.71	0.69	0.52	0.66	
	$^3B_u(V; \pi \rightarrow \pi^*)$	3.36	3.40	3.19	Formamide	$^1A''(V; n \rightarrow \pi^*)$	5.65	5.66	5.45	5.71
	$^3A_g(V; \pi \rightarrow \pi^*)$	5.20	5.32	4.93		$^1A'(R; n \rightarrow 3s)$	6.77	6.80	6.64	6.98
	$^3B_g(R; \pi \rightarrow 3s)$	6.29	6.44	6.26	$^1A'(R; n \rightarrow 3p)$	7.38	7.45	7.32	7.64	
	Carbon trimer	$^1\Delta_g(V; n, n \rightarrow \pi^*, \pi^*)$	5.22	5.08	4.85 <sup>b</sup>	$^1A'(V; \pi \rightarrow \pi^*)$	7.63	7.88	7.13	7.64
$^1\Sigma_g^+(V; n, n \rightarrow \pi^*, \pi^*)$		5.91	5.82	5.58 <sup>c</sup>	$^3A''(V; n \rightarrow \pi^*)$	5.38	5.36	5.16	5.38	
Cyanoacetylene	$^1\Sigma^-(V; \pi \rightarrow \pi^*)$	5.80	5.85	5.47	$^3A'(V; \pi \rightarrow \pi^*)$	5.81	5.88	5.62	5.90	
	$^1\Delta(V; \pi \rightarrow \pi^*)$	6.07	6.13	5.78	Furan	$^1A_2(R; \pi \rightarrow 3s)$	6.09	6.16	6.04	6.28
	$^3\Sigma^+(V; \pi \rightarrow \pi^*)$	4.44	4.45	4.04		$^1B_2(V; \pi \rightarrow \pi^*)$	6.37	6.59	6.02	6.20
	$^3\Delta(V; \pi \rightarrow \pi^*)$	5.21	5.21	4.86	$^1A_1(V; \pi \rightarrow \pi^*)$	6.56	6.66	6.10	6.77	
$^1A''[F](V; \pi \rightarrow \pi^*)$	3.54	3.67	3.47	$^1B_1(R; \pi \rightarrow 3p)$	6.64	6.79	6.63	6.71		
Cyanoformaldehyde	$^1A''(V; n \rightarrow \pi^*)$	3.81	3.98	3.67	$^1A_2(R; \pi \rightarrow 3p)$	6.81	6.87	6.77	6.99	
	$^1A''(V; \pi \rightarrow \pi^*)$	6.46	6.79	6.43	$^1B_2(R; \pi \rightarrow 3p)$	7.24	7.11	6.84	7.01	
	$^3A''(V; n \rightarrow \pi^*)$	3.44	3.46	3.25	$^3B_2(V; \pi \rightarrow \pi^*)$	4.20	4.26	4.01	4.42	
	$^3A'(V; \pi \rightarrow \pi^*)$	5.01	5.25	5.03	$^3A_1(V; \pi \rightarrow \pi^*)$	5.46	5.50	5.09	5.60	
Cyanogen	$^1\Sigma_u^-(V; \pi \rightarrow \pi^*)$	6.39	6.40	6.03	$^3A_2(R; \pi \rightarrow 3s)$	6.02	6.16	5.99	6.08	
	$^1\Delta_u(V; \pi \rightarrow \pi^*)$	6.66	6.70	6.35	$^3B_1(R; \pi \rightarrow 3p)$	6.59	6.76	6.60	6.68	
	$^3\Sigma_u^+(V; \pi \rightarrow \pi^*)$	4.91	4.85	4.46	Glyoxal	$^1A_u(V; n \rightarrow \pi^*)$	2.88	2.82	2.51	2.90
	$^1\Sigma_u^+[F](V; \pi \rightarrow \pi^*)$	5.05	5.07	4.74		$^1B_g(V; n \rightarrow \pi^*)$	4.24	4.21	3.85	4.30
Cyclopentadiene	$^1B_2(V; \pi \rightarrow \pi^*)$	5.56	5.96	5.62	$^1A_g(V; n, n \rightarrow \pi^*, \pi^*)$	5.61	5.37	5.21	5.52	
	$^1A_2(R; \pi \rightarrow 3s)$	5.78	5.88	5.78	$^1B_g(V; n \rightarrow \pi^*)$	6.57	6.52	5.54	6.64	
	$^1B_1(R; \pi \rightarrow 3p)$	6.41	6.59	6.44	$^1B_u(R; n \rightarrow 3p)$	7.71	7.61	7.34	7.84	
	$^1A_2(R; \pi \rightarrow 3p)$	6.46	6.55	6.46	$^3A_u(V; n \rightarrow \pi^*)$	2.49	2.41	2.12	2.49	
	$^1B_2(R; \pi \rightarrow 3p)$	6.56	6.72	6.56	$^3B_g(V; n \rightarrow \pi^*)$	3.89	3.90	3.53	3.99	
	$^1A_1(V; \pi \rightarrow \pi^*)$	6.52	6.63	6.13	$^3B_u(V; \pi \rightarrow \pi^*)$	5.15	5.14	4.91	5.17	
	$^3B_2(V; \pi \rightarrow \pi^*)$	3.31	3.34	3.09	$^3A_g(V; \pi \rightarrow \pi^*)$	6.30	6.32	6.02	6.33	
	$^3A_1(V; \pi \rightarrow \pi^*)$	5.11	5.14	4.78	Imidazole	$^1A''(R; \pi \rightarrow 3s)$	5.70	5.88	5.66	5.93
	$^3A_2(R; \pi \rightarrow 3s)$	5.73	5.91	5.74		$^1A'(R; \pi \rightarrow 3p)$	6.41	6.69	6.45	6.73
	$^3B_1(R; \pi \rightarrow 3p)$	6.36	6.59	6.42	$^1A''(R; \pi \rightarrow 3p)$	6.50	6.57	6.47	6.83	
Cyclopropene	$^1B_1(V; \sigma \rightarrow \pi^*)$	6.68	6.86	6.58	$^1A''(V; n \rightarrow \pi^*)$	6.71	6.94	6.57	6.96	
	$^1B_2(V; \pi \rightarrow \pi^*)$	6.79	6.89	6.47	$^1A'(V; \pi \rightarrow \pi^*)$	6.86	6.88	6.46	7.00	
	$^3B_2(V; \pi \rightarrow \pi^*)$	4.38	4.47	4.27	$^1A'(R; n \rightarrow 3s)$	7.00	7.10	6.91	7.20	
	$^3B_1(V; \sigma \rightarrow \pi^*)$	6.45	6.56	6.32	$^3A'(V; \pi \rightarrow \pi^*)$	4.73	4.78	4.52	4.86	
				$^3A''(R; \pi \rightarrow 3s)$	5.66	5.86	5.63	5.91		
				$^3A'(V; \pi \rightarrow \pi^*)$	5.74	5.85	5.48	5.91		
				$^3A''(V; n \rightarrow \pi^*)$	6.31	6.44	6.10	6.48		

<sup>a</sup>Significant double excitation character; <sup>b</sup>Level shift=0.4; <sup>c</sup>Level shift=0.5.

Table 2: Comparisons between TBEs taken from Table 6 of Ref. 123 and Table 11 of Ref. 125 and VTEs computed at the CASPT2(IPEA), CASPT2(NOIPEA), and (PC-)NEVPT2 values. All values are given in eV and computed with the *aug-cc-pVTZ* basis.

Compound	State	TBE	CASPT2		NEVPT2	Compound	State	TBE	CASPT2		NEVPT2	
			IPEA	NOIPEA				IPEA	NOIPEA			
Isobutene	$^1B_1(R; \pi \rightarrow 3s)$	6.46	6.74	6.58	6.63	Pyridine	$^1B_1(V; n \rightarrow \pi^*)$	4.95	5.15	4.81	5.15	
	$^1A_1(R; \pi \rightarrow 3p)$	7.01	7.32	7.14	7.20		$^1B_2(V; \pi \rightarrow \pi^*)$	5.14	5.18	4.76	5.31	
	$^3A_1(V; \pi \rightarrow \pi^*)$	4.53	4.59	4.41	4.61		$^1A_2(V; n \rightarrow \pi^*)$	5.40	5.46	5.03	5.29	
Ketene	$^1A_2(V; \pi \rightarrow \pi^*)$	3.86	3.92	3.70	3.93	$^1A_1(V; \pi \rightarrow \pi^*)$	6.62	6.82	6.17	6.69		
	$^1B_1(R; \pi \rightarrow 3s)$	6.01	5.99	5.79	6.09	$^1A_1(R; n \rightarrow 3s)$	6.76	6.90	6.67	6.99		
	$^1A_2(R; \pi \rightarrow 3p)$	7.18	7.25	7.05	7.28	$^1A_2(R; \pi \rightarrow 3s)$	6.82	7.08	6.86	6.86		
	$^3A_2(V; \pi \rightarrow \pi^*)$	3.77	3.81	3.59	3.80	$^1B_2(V; \pi \rightarrow \pi^*)$	7.40	7.92	7.67	7.83		
	$^3A_1(V; \pi \rightarrow \pi^*)$	5.61	5.65	5.43	5.65	$^1B_1(R; \pi \rightarrow 3p)$	7.38	7.70	7.51	7.45		
	$^3B_1(R; n \rightarrow 3s)$	5.79	5.79	5.60	5.89	$^1A_1(V; \pi \rightarrow \pi^*)$	7.39	7.66	6.63	6.97		
	$^3A_2(R; \pi \rightarrow 3p)$	7.12	7.21	7.01	7.24	$^3A_1(V; \pi \rightarrow \pi^*)$	4.30	4.39	4.06	4.60		
	$^1A''[F](V; \pi \rightarrow \pi^*)$	1.00	1.05	0.88	1.02	$^3B_1(V; n \rightarrow \pi^*)$	4.46	4.48	4.21	4.58		
	$^1B_2(V; \pi \rightarrow \pi^*)$	4.28	4.40	4.12	4.37	$^3B_2(V; \pi \rightarrow \pi^*)$	4.79	4.86	4.53	4.88		
Methylenecyclopropene	$^1B_1(R; \pi \rightarrow 3s)$	5.44	5.57	5.44	5.49	$^3A_1(V; \pi \rightarrow \pi^*)$	5.04	5.09	4.63	5.19		
	$^1A_2(R; \pi \rightarrow 3p)$	5.96	6.09	5.97	6.00	$^3A_2(V; n \rightarrow \pi^*)$	5.36	5.33	4.96	5.33		
	$^1A_1(V; \pi \rightarrow \pi^*)$	6.12	6.26	6.15	6.36	$^3B_2(V; \pi \rightarrow \pi^*)$	6.24	6.40	5.99	6.29		
	$^3B_2(V; \pi \rightarrow \pi^*)$	3.49	3.57	3.34	3.66	Pyrimidine	$^1B_1(V; n \rightarrow \pi^*)$	4.44	4.44	4.07	4.55	
	$^3A_1(V; \pi \rightarrow \pi^*)$	4.74	4.82	4.58	4.87		$^1A_2(V; n \rightarrow \pi^*)$	4.85	4.80	4.36	4.84	
	$^1A''(V; n \rightarrow \pi^*)$	1.96	1.84	1.60	1.91		$^1B_2(V; \pi \rightarrow \pi^*)$	5.38	5.40	5.01	5.53	
	$^1A'(V; n, n \rightarrow \pi^*, \pi^*)$	4.76	4.69	4.67	4.73		$^1A_2(V; n \rightarrow \pi^*)$	5.92	5.92	5.32	6.02	
	$^1A'(R; n \rightarrow 3s)$	6.29	6.32	6.07	6.38		$^1B_1(V; n \rightarrow \pi^*)$	6.26	6.31	5.65	6.40	
	$^3A''(V; n \rightarrow \pi^*)$	1.16	1.00	0.75	1.08		$^1B_2(R; n \rightarrow 3s)$	6.70	6.85	6.50	6.77	
$^3A'(V; \pi \rightarrow \pi^*)$	5.60	5.52	5.37	5.54	$^1A_1(V; \pi \rightarrow \pi^*)$		6.88	7.23	6.88	7.11		
$^1A''[F](V; n \rightarrow \pi^*)$	1.67	1.55	1.32	1.62	$^3B_1(V; n \rightarrow \pi^*)$		4.09	4.05	3.67	4.17		
$^1A''(V; n \rightarrow \pi^*)$	3.80	3.92	3.64	3.95	$^3A_1(V; \pi \rightarrow \pi^*)$		4.51	4.56	4.25	4.67		
Propynal	$^1A''(V; \pi \rightarrow \pi^*)$	5.54	5.82	5.49	5.50	$^3A_2(V; n \rightarrow \pi^*)$	4.66	4.63	4.16	4.72		
	$^3A''(V; n \rightarrow \pi^*)$	3.47	3.48	3.26	3.59	$^3B_2(V; \pi \rightarrow \pi^*)$	4.96	4.98	4.67	5.01		
	$^3A'(V; \pi \rightarrow \pi^*)$	4.47	4.59	4.30	4.63	Pyrrole	$^1A_2(R; \pi \rightarrow 3s)$	5.24	5.38	5.23	5.51	
	Pyrazine	$^1B_{3u}(V; n \rightarrow \pi^*)$	4.15	4.09	3.66		4.17	$^1B_1(R; \pi \rightarrow 3p)$	6.00	6.26	6.07	6.32
		$^1A_u(V; n \rightarrow \pi^*)$	4.98	4.76	4.26		4.77	$^1A_2(R; \pi \rightarrow 3p)$	6.00	6.16	6.02	6.44
		$^1B_{2u}(V; \pi \rightarrow \pi^*)$	5.02	5.13	4.65		5.32	$^1B_2(V; (\pi \rightarrow \pi^*))$	6.26	6.62	6.36	6.48
		$^1B_{2g}(V; n \rightarrow \pi^*)$	5.71	5.68	5.27		5.88	$^1A_1(V; \pi \rightarrow \pi^*)$	6.30	6.41	5.84	6.53
		$^1A_g(R; n \rightarrow 3s)$	6.65	6.66	6.27		6.70	$^1B_2(R; \pi \rightarrow 3p)$	6.83	6.75	6.11	6.62
		$^1B_{1g}(V; n \rightarrow \pi^*)$	6.74	6.61	6.07		6.75	$^3B_2(V; \pi \rightarrow \pi^*)$	4.51	4.57	4.30	4.74
$^1B_{1u}(V; \pi \rightarrow \pi^*)$		6.88	7.14	6.72	6.81		$^3A_2(R; \pi \rightarrow 3s)$	5.21	5.41	5.21	5.49	
$^1B_{1g}(R; \pi \rightarrow 3s)$		7.21	7.41	7.27	7.33		$^3A_1(V; \pi \rightarrow \pi^*)$	5.45	5.50	5.03	5.56	
$^1B_{2u}(R; n \rightarrow 3p)$		7.24	7.34	6.93	7.25	$^3B_1(R; \pi \rightarrow 3p)$	5.91	6.22	6.03	6.28		
$^1B_{1u}(R; n \rightarrow 3p)$	7.44	7.55	7.08	7.42	Streptocyanine-C1	$^1B_2(V; \pi \rightarrow \pi^*)$	7.13	7.17	6.76	7.13		
$^1B_{1u}(V; \pi \rightarrow \pi^*)$	7.98	8.59	7.96	8.25		$^3B_2(V; \pi \rightarrow \pi^*)$	5.52	5.49	5.22	5.52		
$^3B_{3u}(V; n \rightarrow \pi^*)$	3.59	3.49	3.08	3.56		Tetrazine	$^1B_{3u}(V; n \rightarrow \pi^*)$	2.47	2.31	1.91	2.35	
$^3B_{1u}(V; \pi \rightarrow \pi^*)$	4.35	4.44	4.15	4.57	$^1A_u(V; n \rightarrow \pi^*)$		3.69	3.49	3.00	3.58		
$^3B_{2u}(V; (\pi \rightarrow \pi^*))$	4.39	4.44	4.09	4.42	$^1A_g(V; n, n \rightarrow \pi^*, \pi^*)$		4.61	4.57	4.32	4.61		
$^3A_u(V; n \rightarrow \pi^*)$	4.93	4.73	4.21	4.75	$^1B_{1g}(V; n \rightarrow \pi^*)$		4.93	4.83	4.33	4.95		
$^3B_{2g}(V; n \rightarrow \pi^*)$	5.08	5.04	4.66	5.21	$^1B_{2u}(V; \pi \rightarrow \pi^*)$		5.21	5.31	4.84	5.56		
$^3B_{1u}(V; \pi \rightarrow \pi^*)$	5.28	5.29	4.92	5.35	$^1B_{2g}(V; n \rightarrow \pi^*)$		5.45	5.38	4.90	5.63		
Pyridazine	$^1B_1(V; n \rightarrow \pi^*)$	3.83	3.74	3.36	3.80		$^1A_u(V; n \rightarrow \pi^*)$	5.53	5.51	4.92	5.62	
	$^1A_2(V; n \rightarrow \pi^*)$	4.37	4.29	3.87	4.40		$^1B_{3g}(V; n, n \rightarrow \pi^*, \pi^*)$	6.15	5.85	5.22	6.15	
	$^1A_1(V; \pi \rightarrow \pi^*)$	5.26	5.34	4.87	5.58		$^1B_{2g}(V; n \rightarrow \pi^*)$	6.12	5.96	5.18	6.13	
	$^1A_2(V; n \rightarrow \pi^*)$	5.72	5.73	5.19	5.88	$^1B_{1g}(V; n \rightarrow \pi^*)$	6.91	6.57	5.40	6.76		
	$^1B_2(R; n \rightarrow 3s)$	6.17	6.18	5.90	6.21	$^3B_{3u}(V; n \rightarrow \pi^*)$	1.85	1.70	1.31	1.73		
	$^1B_1(V; n \rightarrow \pi^*)$	6.37	6.50	5.94	6.64	$^3A_u(V; n \rightarrow \pi^*)$	3.45	3.26	2.78	3.36		
	$^1B_2(V; \pi \rightarrow \pi^*)$	6.75	7.21	6.74	7.10	$^3B_{1g}(V; n \rightarrow \pi^*)$	4.20	4.10	3.62	4.24		
	$^3B_1(V; n \rightarrow \pi^*)$	3.19	3.08	2.72	3.13	$^3B_{1u}(V; \pi \rightarrow \pi^*)$	4.49	4.55	4.29	4.70		
	$^3A_2(V; n \rightarrow \pi^*)$	4.11	4.01	3.59	4.14	$^3B_{2u}(V; \pi \rightarrow \pi^*)$	4.52	4.55	4.20	4.58		
$^3B_2(V; \pi \rightarrow \pi^*)$	4.34	4.40	4.12	4.49	$^3B_{2g}(V; n \rightarrow \pi^*)$	5.04	5.02	4.53	5.27			
$^3A_1(V; \pi \rightarrow \pi^*)$	4.82	4.87	4.47	4.94	$^3A_u(V; n \rightarrow \pi^*)$	5.11	5.07	4.43	5.13			
					$^3B_{3g}(V; n, n \rightarrow \pi^*, \pi^*)$	5.51	5.39	4.86	5.51			
					$^3B_{1u}(V; \pi \rightarrow \pi^*)$	5.42	5.46	5.08	5.56			

Table 3: Comparisons between TBEs taken from Table 6 of Ref. 123 and Table 11 of Ref. 125 and VTEs computed at the CASPT2(IPEA), CASPT2(NOIPEA), and (PC-)NEVPT2 values. All values are given in eV and computed with the *aug-cc-pVTZ* basis.

Compound	State	TBE	CASPT2	NEVPT2	Compound	State	TBE	CASPT2	NEVPT2			
		IPEA NOIPEA					IPEA NOIPEA					
Thioacetone	$^1A_2(V; n \rightarrow \pi^*)$	2.53	2.58	2.33	Thiopropynal	$^1A''(V; n \rightarrow \pi^*)$	2.03	2.05	1.84	2.05		
	$^1B_2(R; n \rightarrow 4s)$	5.56	5.60	5.48		$^3A''(V; n \rightarrow \pi^*)$	1.80	1.81	1.59	1.81		
	$^1A_1(V; \pi \rightarrow \pi^*)$	5.88	6.66	6.41	6.24	Triazine	$^1A_1''(V; n \rightarrow \pi^*)$	4.72	4.62	3.90	4.61	
	$^1B_2(R; n \rightarrow 4p)$	6.51	6.51	6.40	6.62		$^1A_2''(V; n \rightarrow \pi^*)$	4.75	4.77	4.39	4.89	
	$^1A_1(R; n \rightarrow 4p)$	6.61	6.42	5.98	6.52		$^1E''(V; n \rightarrow \pi^*)$	4.78	4.76	4.14	4.88	
	$^3A_2(V; n \rightarrow \pi^*)$	2.33	2.34	2.09	2.32		$^1A_2'(V; \pi \rightarrow \pi^*)$	5.75	5.76	5.32	5.95	
	$^3A_1(V; \pi \rightarrow \pi^*)$	3.45	3.48	3.29	3.48		$^1A_1'(V; \pi \rightarrow \pi^*)$	7.24	7.43	6.89	7.30	
	Thiophene	$^1A_1(V; \pi \rightarrow \pi^*)$	5.64	5.84	5.21		5.84	$^1E'(R; n \rightarrow 3s)$	7.32	7.48	7.15	7.45
		$^1B_2(V; \pi \rightarrow \pi^*)$	5.98	6.35	5.89		6.10	$^1E''(V; n \rightarrow \pi^*)$	7.78	7.75	7.04	7.98
$^1A_2(R; \pi \rightarrow 3s)$		6.14	6.28	6.07	6.20	$^1E'(V; \pi \rightarrow \pi^*)$	7.94	8.65	7.69	8.34		
$^1B_1(R; \pi \rightarrow 3p)$		6.14	6.21	5.90	6.19	$^3A_2''(V; n \rightarrow \pi^*)$	4.33	4.37	3.99	4.51		
$^1A_2(R; \pi \rightarrow 3p)$		6.21	6.32	5.98	6.40	$^3E''(V; n \rightarrow \pi^*)$	4.51	4.47	3.88	4.61		
$^1B_1(R; \pi \rightarrow 3s)$		6.49	6.54	6.22	6.71	$^3A_1''(V; n \rightarrow \pi^*)$	4.73	4.70	3.94	4.71		
$^1B_2(R; \pi \rightarrow 3p)$		7.29	7.29	7.03	7.25	$^3A_1'(V; \pi \rightarrow \pi^*)$	4.85	4.88	4.54	5.05		
$^1A_1(V; \pi \rightarrow \pi^*)$		7.31	7.62	6.85	7.39	$^3E'(V; \pi \rightarrow \pi^*)$	5.59	5.62	5.20	5.73		
$^3B_2(V; \pi \rightarrow \pi^*)$		3.92	3.98	3.71	4.13	$^3A_2'(V; \pi \rightarrow \pi^*)$	6.62	6.62	6.12	6.36		
$^3A_1(V; \pi \rightarrow \pi^*)$		4.76	4.85	4.39	4.84							
$^3B_1(R; \pi \rightarrow 3p)$	5.93	5.97	5.64	5.98								
$^3A_2(R; \pi \rightarrow 3s)$	6.08	6.22	6.01	6.14								

out IPEA shift. The VTEs of the valence singlet and triplet  $\pi \rightarrow \pi^*$  obtained from CASPT2(IPEA) and NEVPT2 levels of theory are also close from one another, with the NEVPT2 values being in perfect agreement with the TBEs.<sup>123</sup> The absence of IPEA correction again yields significantly underestimated values.

### 3.2 Molecules with four non-hydrogen atoms

For these 15 systems, as well as for the larger systems treated below, the description of the corresponding active spaces can be found in our earlier contribution.<sup>125</sup>

*Acetone and thioacetone.* Previous works reporting CASPT2 and NEVPT2 VTEs exist for acetone.<sup>35,48,53,125,150,155</sup> The present CASPT2(IPEA) results are in excellent agreement with both their counterparts of Ref. 155 and the TBEs, except for the  $^1A_2(n \rightarrow 3p)$  transition that appears significantly too high. Interestingly, the VTE associated with this transition is nicely reproduced by CASPT2(NOIPEA). For the lowest  $n \rightarrow \pi^*$  excitations of both spin symmetries, CASPT2(IPEA) provides VTEs in very close agreement with the ones reported by Roos and co-workers.<sup>35</sup> The NEVPT2 estimates are accurate for the valence ESs, but too large for the Rydberg ESs. In thioacetone, the valence  $^1A_1(\pi \rightarrow \pi^*)$  transitions is considerably overshoot by all three models, NEVPT2 being the most accurate with an error of 0.36 eV. For all other ESs, both CASPT2(IPEA) and NEVPT2 deliver very satisfying accuracy, with no absolute error exceeding 0.2 eV. Again, turning off the IPEA shift is detrimental for most transitions, especially the valence ones.

*Acrolein.* The comprehensive previous CASPT2 study of acrolein is likely the work of Aquilante, Barone, and Roos,<sup>156</sup> which reports many transitions. For the nine ESs listed in Table 1, the lack of IPEA shift clearly yields significant underestimations (except for the ES having a significant share of double excitation character) whereas both CASPT2(IPEA) and NEVPT2 values are trustworthy, the latter leading a small average deviations as compared to the TBEs. However, one specific challenging ES is the second  $^1A''(n \rightarrow \pi^*)$  valence

excitation, for which the VTEs produced by all approaches seem inconsistent. One should note, however, that this transition has a significant contribution from the doubly-excited configurations,<sup>125</sup> making its CCSDT-based TBE likely less accurate.

*Butadiene.* The relative VTEs of the optically bright  $^1B_u$  and dark  $^2A_g$  ESs have certainly been the topic of many theoretical studies, given both the experimental interest and the mixed single/double excitation character of the latter ES.<sup>31,125,157–167</sup> The first very reliable estimates are likely due to Watson and Chan,<sup>159</sup> who showed that the bright ESs should be slightly lower in energy. Our 6.22 and 6.50 eV TBEs, based on CCSDTQ and FCI results, respectively,<sup>124,125</sup> follow this trend and can likely be considered as reliable. As can be seen, none of the three tested multiconfigurational approaches accurately reproduced the gap between these two ESs, CASPT2(NOIPEA) being very poor, confirming that butadiene remains a particularly stringent test. For the other ESs, the trends noted above are conserved, i.e., both NEVPT2 and CASPT2(IPEA) provide reliable valence/overestimated Rydberg VTEs, whereas CASPT2(NOIPEA) tends to be more reliable for the Rydberg ESs.

*Cyanoacetylene, cyanogen, and diacetylene.* For these three closely related linear molecules, the TBEs are based on CCSDTQ, and the difference between CC3, CCSDT, and CCSDTQ values are totally negligible,<sup>125</sup> strongly hinting that the TBEs are highly trustworthy. On average, both CASPT2(IPEA) and NEVPT2 are rather competitive with deviations no larger than  $\pm 0.10$  eV except for the fluorescence of cyanoacetylene with CASPT2(IPEA). In contrast, CASPT2(NOIPEA) systematically produces underestimated VTEs for all 13 ESs of this series.

*Cyanoformaldehyde, propynal, and thiopropynal.* These three molecules possess a conjugated carbonyl group and are of  $C_s$  symmetry. All the considered ESs have a strong single-excitation character.<sup>125</sup> The VTEs of the lowest singlet transitions, of  $n \rightarrow \pi^*$  nature, are almost identical between



NEVPT2 and CASPT2(IPEA), with absolute values slightly exceeding the TBEs. Surprisingly, the differences are larger for the corresponding triplet ES in both cyanoformaldehyde and propynal, for which CASPT2(IPEA) is extremely accurate. In contrast, for the  $A'$  ( $\pi \rightarrow \pi^*$ ) transitions, not applying an IPEA shift seems beneficial in the case of cyanoformaldehyde.

*Cyclopropenone, cyclopropenethione, and methylenecyclopropene.* These three molecules are characterized by a three-membered  $sp^2$  carbon cycle conjugated to an external  $\pi$  bond. There are previous multiconfigurational studies for these three compounds.<sup>42,125,145,168,169</sup> For cyclopropenone, the VTEs of the singlet and triplet valence  $n \rightarrow \pi^*$  transitions of  $A_2$  symmetry, singlet  $n \rightarrow 3s$  Rydberg transition of  $B_2$  symmetry, and triplet  $\pi \rightarrow \pi^*$  transition of  $A_1$  symmetry obtained with CASPT2(IPEA) nicely fit the reference CCSDTQ-based values.<sup>125</sup> In contrast, the VTEs of the singlet and triplet  $n \rightarrow \pi^*$  transitions of  $B_1$  symmetry (singlet  $\pi \rightarrow \pi^*$  transition of  $B_2$  symmetry) are underestimated (overestimated) by all second-order approaches assessed here. For cyclopropenethione, the CASPT2 values reported by Serrano-Andrés et al.<sup>42</sup> are typically in between the CASPT2(IPEA) and CASPT2(NOIPEA) values listed in Table 1. CASPT2(IPEA) outperforms the two other methods reported in this Table, except for the  ${}^1B_2(n \rightarrow 3s)$  and  ${}^1A_1(\pi \rightarrow \pi^*)$  transitions. For methylenecyclopropene, the current CASPT2(NOIPEA) results fit well with those reported by Roos and co-workers.<sup>168</sup> Pinpointing the most suitable level of theory is challenging from the data of Table 2 as the MAEs are similar: 0.11, 0.09, and 0.12 eV for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively.

*Glyoxal.* In glyoxal, one needs to separate the lowest  ${}^1A_g$  ES, of pure double ( $n, n \rightarrow \pi^*, \pi^*$ ) nature, from all other considered ESs that are strongly dominated by single excitations. For the latter set of transitions, both CASPT2(IPEA) and NEVPT2 typically provide accurate estimates, whereas the lack of IPEA correction gives significant underestimations, even for the Rydberg transition considered herein. For the specific  ${}^1A_g$  ES, our TBE is based on FCI,<sup>124,125</sup> and the NEVPT2 VTE is compatible with this estimate, whereas both CASPT2 approaches deliver too small transition energies. Interestingly the SAC-CI method in its so-called general- $R$  form provides a VTE of 5.66 eV for this state,<sup>158</sup> actually outperforming the three multiconfigurational methods considered in the present study.

*Isobutene.* Similarly to other cases, both CASPT2(IPEA) and NEVPT2 significantly overshoot the singlet TBEs of the Rydberg ESs, whereas for the lowest triplet of valence nature, these two approaches yield results within 0.1 eV of the reference value.

### 3.3 Five-membered cycles

*Cyclopentadiene.* There are several previous CASPT2<sup>32,48,53,145</sup> and NEVPT2<sup>125</sup> studies of VTEs in cyclopentadiene. None of the ES treated here, except the  ${}^1A_1(\pi \rightarrow \pi^*)$  transition, has a significant contribution from the doubly-excited determinants. However, even for this transition, one notes in Table 1 a reasonable agreement between the CASPT2(IPEA), NEVPT2, and the CCSDT-based TBEs (that all slightly ex-

ceed earlier estimates<sup>32,48,53</sup>) as well as the most recent experimental value we are aware of.<sup>170</sup> For all other singlet transitions, including the lowest  ${}^1B_2(\pi \rightarrow \pi^*)$  excitation, CASPT2(NOIPEA) provides VTEs closer to the TBEs than CASPT2(IPEA). For the triplet ESs considered here, one finds again the usual pattern with valence (Rydberg) VTEs more accurate when the IPEA shift is turned on (off). Considering the 10 ESs, one gets MAEs of 0.15, 0.11, and 0.10 eV for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively.

*Furan.* Unsurprisingly, furan was also investigated with several multireference methods.<sup>32,48,53,145,171,172</sup> We consider only ESs possessing a highly-dominant single excitation character. For both the singlet and triplet valence  $\pi \rightarrow \pi^*$  transitions of  $B_2$  and  $A_1$  symmetries, the present CASPT2(NOIPEA) values are reasonably close to those determined by Roos' group,<sup>32</sup> whereas blueshifts are found when including the IPEA shift.<sup>48,53</sup> An interesting case is the  ${}^1B_2(n \rightarrow 3p)$  excitation for which our CCSDT-based TBE of 7.24 eV is larger than the estimates given in most previous theoretical studies (ca. 6.5–6.9 eV).<sup>32,48,53,171–173</sup> For this particular ES, the present NEVPT2 and CASPT2(IPEA) estimates are also above 7 eV, indicating that this ES is very sensitive to the employed methodological details. For the 10 transitions reported in Table 1, the MAEs are 0.11, 0.19, and 0.16 eV for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively, that is, the relative ranking of the methods differs from the one obtained for cyclopentadiene.

*Imidazole.* The most detailed previous CASPT2 investigation is due to Serrano-Andrés and co-workers,<sup>37</sup> and the VTEs we report here with CASPT2(NOIPEA) for the Rydberg transitions, as well as the singlet and triplet valence  $\pi \rightarrow \pi^*$  transitions of  $A'$  symmetry are quite close to these previous values. The quality of these VTEs, as compared to the TBEs,<sup>125</sup> are rather contrasted.

*Pyrrrole.* The VTEs of pyrrole were extensively studied using second-order perturbation theory.<sup>32,41,48,53,106,125,173</sup> We underline that our TBEs<sup>125</sup> are in very nice agreement with XMS-CASPT2 results obtained by the González group,<sup>106</sup> while the present CASPT2(NOIPEA) values are larger than the 1993 VTEs proposed by Serrano-Andrés et al.<sup>32</sup> For the singlet transitions, it is difficult to highlight clear trends for both CASPT2 approaches, whereas for the triplet ESs the usual tendencies are found. Interestingly NEVPT2 systematically overestimates the VTEs for all considered ESs, except for the highest-lying Rydberg state.

*Thiophene.* The optical properties of the most popular sulfur-bearing cycle have been investigated both experimentally,<sup>174,175</sup> and with multiconfigurational approaches.<sup>33,125,176</sup> Again, all ESs listed in Table 3 are of dominant single-excitation character though the two singlet  $A_1$  transitions have non-negligible contributions from higher excitations. CASPT2(NOIPEA) systematically yields too small VTEs, with an average deviation of  $-0.25$  eV. Turning on the IPEA shift somehow overcorrects, as CASPT2(IPEA) produces systematic overestimations by an average of  $+0.13$  eV. NEVPT2 delivers a more balanced treatment, with no deviation exceeding a quarter of an eV and a MAE of 0.11 eV.



### 3.4 Six-membered cycles

*Benzene.* Previous multiconfigurational studies,<sup>48,53,124,125,177,178</sup> as well as refined experimental measurements<sup>179,180</sup> are again available for benzene. We find that our CASPT2(NOIPEA) data for the singlet Rydberg transitions are systematically higher in energy than the previous CASPT2 results of Roos and co-workers,<sup>177</sup> whereas our TBEs reasonably fit published RASPT2 data.<sup>178</sup> The trends obtained by analyzing the results of Table 1 are rather clear: (i) CASPT2(IPEA) outperforms both CASPT2(NOIPEA) and NEVPT2 for estimating the valence  $\pi \rightarrow \pi^*$  transitions; (ii) CASPT2(NOIPEA) is accurate for Rydberg transitions but leads to small values for the valence ESs.

*Pyrazine.* There are previous multireference estimates of the VTEs of pyrazine,<sup>48,53,124,125,181–184</sup> and all ESs considered can be described as single excitations. Indeed, the ES with the largest contribution from higher excitations, the  $^1B_{1g}$   $n \rightarrow \pi^*$  transition, has a single character of 84.2% according to CC3.<sup>125</sup> From the data listed in Table 2, it is clear that CASPT2(NOIPEA) is not very accurate, but it is not straightforward to determine if CASPT2(IPEA) is superior to NEVPT2 (or the opposite). An interesting specificity is that the lowest  $^1A_u$  ES seems rather challenging for both methods.

*Pyridazine.* As for previous systems, the present VTEs are consistent with previous CASPT2 studies including or not IPEA shifts.<sup>48,53,181,185</sup> It is also noteworthy that our CASPT2(NOIPEA) data are in close proximity with the very recent electron energy-loss experiments,<sup>186</sup> but, as always, such comparison can only be viewed as qualitative. Considering all pyridazine ESs of Table 2, one obtains MAEs of 0.07, 0.44, and 0.12 eV, for CASPT2(IPEA), CASPT2(NOIPEA), and NEVPT2, respectively. The absence of IPEA shift being systematically detrimental, even for the Rydberg  $^1B_2$  transition.

*Pyridine.* The ESs of pyridine were the subject of multiconfigurational<sup>48,53,125,181,184,187</sup> and refined experimental<sup>188,189</sup> studies. We found a good correlation between the current CASPT2(NOIPEA) data and the early estimates of the Roos group.<sup>187</sup> Again, the VTEs of the Rydberg transitions, e.g., the singlet  $\pi/n \rightarrow 3s$  transitions of  $A_1$  and  $A_2$  symmetry, obtained with CASPT2(NOIPEA) are in good agreement with the TBEs, whereas for the valence transitions, both CASPT2(IPEA) and NEVPT2 provide more accurate VTEs. Those two latter methods provide similar MAEs (ca. 0.15 eV) for pyridine.

*Pyrimidine.* Earlier CASPT2 and NEVPT2 estimates are available for this azabenzene.<sup>48,53,125,181,184,190,191</sup> The results of Table 2 reveal the excellent accuracy obtained with CASPT2(IPEA). There is only one ES, namely,  $^1A_1(\pi \rightarrow \pi^*)$ , for which a deviation larger than 0.15 eV could be detected. For this state, of strong single-excitation character,<sup>125</sup> not applying an IPEA shift is beneficial. NEVPT2 also provides very reasonable estimates except for a small overestimation trend. Nevertheless, NEVPT2 outperforms both CASPT2 schemes for the Rydberg  $^1B_2(n \rightarrow 3s)$  transition, which is the unusual outcome for the panel of compounds treated here.

*Tetrazine.* In this intensively studied,<sup>48,53,124,125,192–194</sup>

highly symmetric compound, one can clearly distinguish ESs of single-excitation character to those that are pure double excitations, namely,  $^1A_g$ ,  $^1B_{3g}$ , and  $^3B_{3g}$ .<sup>124,125</sup> For these three particular states, our TBEs were previously classified as “unsafe”,<sup>125</sup> as FCI estimates are beyond computational reach, so that indisputable TBEs are not available. For the  $^1A_g$  transition, the CASPT2(IPEA) and NEVPT2 values are very close from another, but differences of 0.30 and 0.12 eV are obtained for the singlet and triplet  $B_{3g}$  transitions, respectively, justifying their “unsafe” status. Focussing on the single-excitation transitions, one notes very good performances of both CASPT2(IPEA) and NEVPT2 with respective MAEs of 0.11 and 0.12 eV.

*Triazine.* Finally, for the  $D_{3h}$  azabenzene, the present CASPT2(NOIPEA) results are in line with the early investigations of Roos and co-workers<sup>181</sup> and Serrano-Andrés and co-workers<sup>195</sup> whereas the CASPT2(IPEA) data fit reasonably well the CASPT2 values determined later by Thiel’s group.<sup>48,53</sup> When comparing to the TBEs, it is pretty obvious that CASPT2(IPEA) is the most adequate approach for most ESs, the absence of IPEA shift leading to systematic underestimations of the VTEs, whereas NEVPT2 produces slightly too large transition energies.

### 3.5 Statistical analysis

To perform our statistical analysis, we only take into account the VTEs qualified as “safe” in our previous studies.<sup>125,128</sup> Therefore, we removed, for example, transitions of double excitation character for which FCI energies could not be computed, the strongly mixed ESs in formamide, and the other troublemakers. The corresponding statistical indicators are gathered in Table 4 and the histogram representation of the spread of the errors can also be found in Figure 2.

Table 4: Statistical analysis, taking the TBEs as reference, for the four multiconfigurational approaches. At the bottom of the Table, we provide similar data extracted from the QUEST database for single-reference methods.<sup>128</sup> All values are in eV.

Method	Count	MSE	MAE	RMSE	SDE	Max(+)	Max(-)
SA-CASSCF	265	0.12	0.47	0.59	0.56	2.14	-1.18
CASPT2(IPEA)	265	0.06	0.11	0.16	0.15	0.78	-0.44
CASPT2(NOIPEA)	265	-0.26	0.28	0.35	0.24	0.53	-1.51
SC-NEVPT2	265	0.13	0.15	0.19	0.13	0.65	-0.38
PC-NEVPT2	265	0.09	0.13	0.16	0.13	0.46	-0.42
CIS(D)	257	0.16	0.22	0.28	0.23	0.96	-0.69
ADC(2)	252	0.00	0.14	0.19	0.18	0.64	-0.73
CC2	258	0.03	0.15	0.20	0.19	0.59	-0.68
SCS-CC2	258	0.15	0.17	0.21	0.15	0.76	-0.92
CCSD	259	0.10	0.12	0.16	0.05	0.62	-0.17
CC3	262	0.00	0.02	0.03	0.03	0.21	-0.09

In line with the analysis performed for individual molecules, it turns out that CASPT2 delivers slightly too large VTEs when the IPEA shift is turned on (+0.06 eV) but large underestimations when the IPEA shift is set to zero (-0.26 eV). The MAE obtained with CASPT2(IPEA) is 0.11 eV, which can certainly be viewed as acceptable for many practical applications. It, nevertheless, remains far from chemical accuracy (1 kcal.mol<sup>-1</sup> or 0.04 eV). The dispersion of the errors, as measured by the SDE, is also much larger when the IPEA shift is not applied (0.24 eV vs 0.15 eV), which

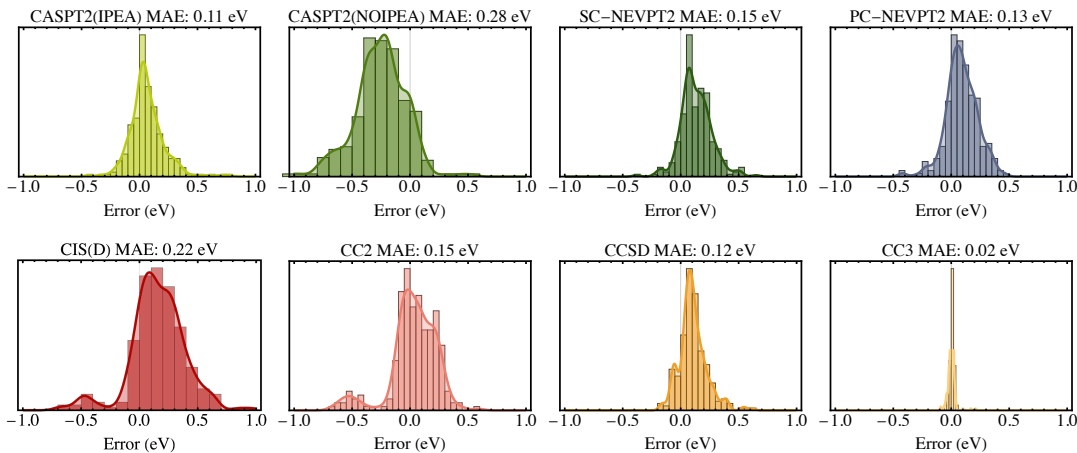


Figure 2: Histograms of the errors obtained for the four multiconfigurational second-order perturbation theory methods and comparisons with a selection of single-reference methods.

clearly shows that the IPEA shift must be applied in practical CASPT2 calculations of VTEs in organic molecules. This observation is consistent with the findings of the González group, who concluded that, for triple- $\zeta$  basis sets, the IPEA shift was beneficial.<sup>109</sup> Interestingly, PC-NEVPT2, which is free of the IPEA dilemma, delivers almost the same performance as CASPT2(IPEA), with slightly larger overestimations of the VTEs, but slightly smaller spread of the errors and maximal deviations. In other words, the present benchmark does not reveal any significant difference in terms of accuracy between these two multiconfigurational approaches. SC-NEVPT2 is found to be only slightly worse than PC-NEVPT2. In contrast (and as expected), SA-CASSCF is unsatisfying with a MAE as large as 0.47 eV and a large dispersion as well.

At the bottom of Table 4, we provide the results obtained for the same set of compounds and ESs with a selection of popular single-reference methods. For the sake of consistency, these values have been straightforwardly extracted from the QUEST database,<sup>128</sup> but are in line with many other benchmarks.<sup>48,70,151,196–198</sup> As can be seen ADC(2), CC2, and CCSD do deliver average deviations of the same order of magnitude as both CASPT2(IPEA) and PC-NEVPT2, but with a slightly larger dispersion of the errors for ADC(2) and CC2. The latter effect can be attenuated by using the spin-scaled variant (SCS) of CC2, but at the cost of inducing a nearly systematic overestimation, as evidenced by the large positive MSE. This means that one should be very cautious in using CASPT2 data as reference to benchmark the quality of ADC(2)’s or CC2’s VTEs in well-behaved transitions. Indeed, the superiority of the former model over the two others is not perfectly clear.

In Table 5, we provide the MAEs determined for various subsets of transitions, whereas Figure 3 provides a graphical comparison for selected families of transitions. Clearly CASPT2(NOIPEA) delivers reasonable estimates for the Rydberg transitions only, which is fully consistent with the analyses made for individual systems above. SA-CASSCF appears more accurate for the triplet, valence, and  $n \rightarrow \pi^*$  transitions than for the singlet, Rydberg, and  $\pi \rightarrow \pi^*$  coun-

terparts, but the deviations remain very large in all cases. Although one can notice small differences for various subsets, i.e., slightly improved performances for triplet, valence, and  $n \rightarrow \pi^*$  ESs as compared to the singlet, Rydberg, and  $\pi \rightarrow \pi^*$  subsets, it is a pleasant outcome that both CASPT2(IPEA) and the two NEVPT2 variants deliver rather equivalent and satisfactory levels of accuracy for all single-excitation subsets. For the systems considered here, rather similar trends were observed for ADC(2) and CC2, though these two models are more effective for the  $n \rightarrow \pi^*$  than the  $\pi \rightarrow \pi^*$  excitations. Finally, the transitions with a dominant contribution from the doubly-excited states are more accurately modeled by NEVPT2 than CASPT2, while single-reference approaches are simply unable to describe these transitions.

Besides, by looking into the evolution of the errors as the size of the compounds increases, one notes a deterioration of the MAEs for CASPT2(IPEA) and both NEVPT2 schemes when modeling larger molecules, with, e.g., MAEs of 0.08, 0.12, and 0.15 eV for the 3, 4, and 5-6 non-hydrogen atom molecules for PC-NEVPT2. The CASPT2(NOIPEA) deviations follow a similar pattern, but with larger errors, with a MAE of 0.20 eV for the 3 non-hydrogen atom molecules, but 0.33 eV for the 5-6 non-hydrogen atom compounds.

We can compare the deviations obtained here to the ones reported in Ref. 68 using Thiel’s CC3/TZVP values as reference. For 121 singlet (72 triplets) valence ESs, Schapiro *et al.* obtained MAEs of 0.23 (0.23), 0.28 (0.25), and 0.21 (0.14) eV for SC-NEVPT2, PC-NEVPT2, and SS-CASPT2(IPEA), respectively, that is, deviations significantly larger than the present ones. However, we note that their key conclusions stating that the three approaches deliver similar results, and that CASPT2 is slightly more accurate for the triplet ESs are fully consistent with the present findings. Analyzing the data of this earlier work, it appears that differences with the current results can be mostly ascribed to the  $\pi \rightarrow \pi^*$  transitions (that are more accurate here quite possibly due to the use of a larger diffuse-containing basis set), and likely to the total absence of Rydberg and doubly-excited transitions in Thiel’s set. In Ref. 68, the absence of IPEA correction produced

Table 5: MAE determined for several subsets computed at various levels of theory. All values are in eV.

Method	Singlet	Triplet	Valence	Rydberg	$n \rightarrow \pi^*$	$\pi \rightarrow \pi^*$	Double
SA-CASSCF	0.55	0.34	0.44	0.53	0.44	0.44	0.16
CASPT2(IPEA)	0.14	0.07	0.11	0.13	0.08	0.13	0.14
CASPT2(NOIPEA)	0.28	0.29	0.34	0.15	0.45	0.27	0.30
SC-NEVPT2	0.16	0.13	0.15	0.15	0.12	0.17	0.04
PC-NEVPT2	0.14	0.11	0.12	0.15	0.10	0.13	0.06
CIS(D)	0.21	0.24	0.25	0.16	0.21	0.27	
ADC(2)	0.15	0.13	0.13	0.16	0.09	0.16	
CC2	0.16	0.14	0.14	0.17	0.07	0.19	
SCS-CC2	0.16	0.19	0.20	0.10	0.22	0.19	
CCSD	0.15	0.09	0.14	0.09	0.18	0.11	
CC3	0.02	0.01	0.02	0.01	0.01	0.02	

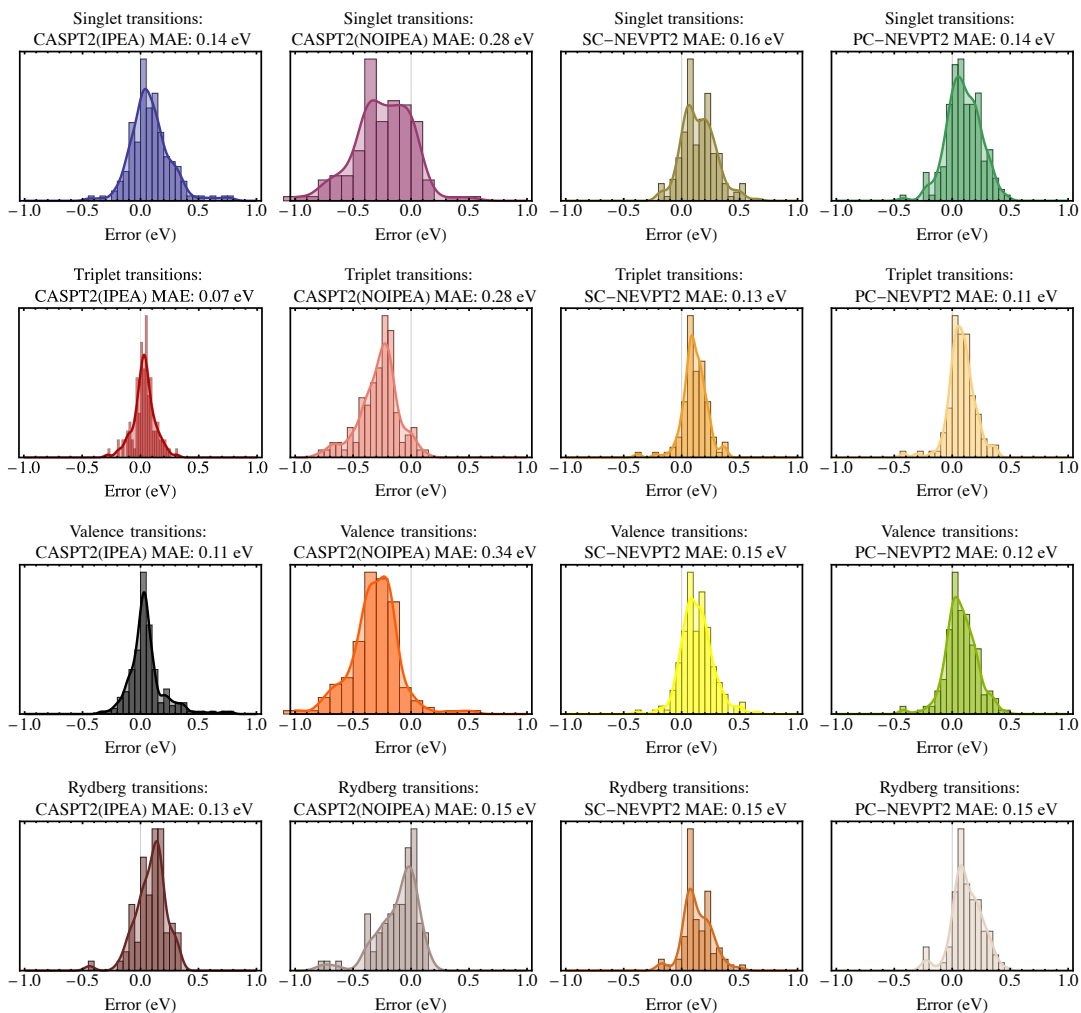


Figure 3: Histograms of the errors for various subsets obtained with the four multiconfigurational second-order perturbation theory methods.

very large underestimations of the CASPT2 VTEs (MSE of  $-0.48$  eV for singlet transitions). This is again consistent with the trend reported here, but with a more dramatic error likely due to the particular focus on valence transitions. Finally, we note that for SA-CASSCF, much larger errors have been reported when using the 122 Thiel CC3/TZVP reference values,<sup>78</sup> with a MAE of 0.98 eV. Nonetheless, the fact that SA-CASSCF is more trustworthy for  $n \rightarrow \pi^*$  than  $\pi \rightarrow \pi^*$  transitions is consistent with this study.<sup>78</sup>

#### 4. CONCLUSIONS

We have computed VTEs of more than 250 ESs in 35 small to medium size organic molecules containing three to six non-hydrogen atoms using four different perturbative approaches and compared these VTEs with previous TBEs established with highly-accurate methods.<sup>125,128</sup>

Besides containing a wealth of data for individual molecules and specific transitions, this study allows to extract general trends. In this regard, the global statistical analyses reveal that the overall performance of CASPT2(NOIPEA) is not up to the mark, except maybe for Rydberg transitions. Indeed, not applying an IPEA shift typically leads to strong underestimations of the transition energies together with a significant spread of the errors. The same statistical analyses indicate that both CASPT2(IPEA) and PC-NEVPT2 do yield small overestimations of the VTEs and provide more accurate estimates for the different subsets of ESs, with global MAEs of 0.11 and 0.13 eV, and SDEs of 0.15 and 0.13 eV, respectively. Neglecting dynamic correlation, as in SA-CASSCF, is clearly very detrimental, whereas using the SC instead of the PC variant of NEVPT2 only slightly deteriorates the results. As compared to single-reference models, one notes that CASPT2(IPEA) and PC-NEVPT2 deliver similar accuracies as ADC(2), CC2, and CCSD, but with smaller dispersion than the first two methods. As a consequence, one should likely be very cautious when comparing ADC(2) and CASPT2 to attribute the error mainly to one of the two methods. It is also noteworthy that the CC3 VTEs remain much more accurate than the multiconfigurational ones.

Of course, the present study has some limitations. First, for each of the ESs treated herein, one can certainly define a larger active space leading to a more accurate VTE. Yet, we trust that our active spaces are very reasonable, that is, they would be aligned with typical choices made by CASSCF experts. Second, though we tested the use of IPEA (0.25 au) or its complete neglect (as typically done in the literature), one could also consider searching for an optimal value. This aspect was previously treated by others<sup>109</sup> and it was demonstrated to be fairly challenging to define a general value applicable across various basis sets and systems. Third (and we wish to particularly stress this point), the present investigation was focussed on ESs with a dominant single-excitation character, meaning that the present conclusions regarding the relative performances of CASSCF, CASPT2, NEVPT2, and the various other single-reference approaches are obviously limited to this specific class of ESs.

#### ACKNOWLEDGEMENTS

RS and DJ are indebted to the Région des Pays de la Loire program for support in the framework of the Opt-Basis grant.

PFL thanks the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation programme (Grant agreement No. 863481) for financial support. This research used computational resources of (i) the GENCI-TGCC (Grant No. 2019-A0060801738); (ii) CALMIP under allocation 2020-18005 (Toulouse); (iii) CCIPL (*Centre de Calcul Intensif des Pays de Loire*); (iv) a local Troy cluster and (v) HPC resources from ArronaxPlus (grant ANR-11-EQPX-0004 funded by the French National Agency for Research).

#### SUPPORTING INFORMATION AVAILABLE

Description of active spaces for molecules not treated previously. Corrections and additions of previous data. CASSCF and SC-NEVPT2 transitions energies. The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/doi>.

#### REFERENCES

- (1) Møller, C.; Plesset, M. S. Note on an Approximation Treatment for Many-Electron Systems. *Phys. Rev.* **1934**, *46*, 618–622.
- (2) Andersson, K.; Malmqvist, P. A.; Roos, B. O.; Sadlej, A. J.; Wolinski, K. Second-Order Perturbation Theory With a CASSCF Reference Function. *J. Phys. Chem.* **1990**, *94*, 5483–5488.
- (3) Andersson, K.; Malmqvist, P.-A.; Roos, B. O. Second-Order Perturbation Theory With a Complete Active Space Self-Consistent Field Reference Function. *J. Chem. Phys.* **1992**, *96*, 1218–1226.
- (4) Hirao, K. Multireference Møller–Plesset Method. *Chem. Phys. Lett.* **1992**, *190*, 374–380.
- (5) Angeli, C.; Cimiraglia, R.; Evangelisti, S.; Leininger, T.; Malrieu, J.-P. Introduction of  $N$ -Electron Valence States for Multireference Perturbation Theory. *J. Chem. Phys.* **2001**, *114*, 10252–10264.
- (6) Angeli, C.; Cimiraglia, R.; Malrieu, J.-P.  $N$ -Electron Valence State Perturbation Theory: A Fast Implementation of the Strongly Contracted Variant. *Chem. Phys. Lett.* **2001**, *350*, 297–305.
- (7) Angeli, C.; Cimiraglia, R.; Malrieu, J.-P.  $N$ -Electron Valence State Perturbation Theory: A Spinless Formulation and an Efficient Implementation of the Strongly Contracted and of the Partially Contracted Variants. *J. Chem. Phys.* **2002**, *117*, 9138–9153.
- (8) Andersson, K.; Roos, B.; Malmqvist, P. A.; Widmark, P.-O. The  $\text{Cr}_2$  Potential Energy Curve Studied with Multiconfigurational Second-Order Perturbation Theory. *Chem. Phys. Lett.* **1994**, *230*, 391–397.
- (9) Andersson, K. The Electronic Spectrum of  $\text{Cr}_2$ . *Chem. Phys. Lett.* **1995**, *237*, 212–221.
- (10) Roos, B. O.; Fülcher, M.; Malmqvist, P.-Å.; Merchán, M.; Serrano-Andrés, L. *Quantum Mechanical Electronic Structure Calculations with Chemical Accuracy*; Springer Netherlands: Dordrecht, 1995; pp 357–438.



- (11) Roos, B. O.; Andersson, K. Multiconfigurational Perturbation Theory with Level Shift — the Cr<sub>2</sub> Potential Revisited. *Chem. Phys. Lett.* **1995**, *245*, 215–223.
- (12) Forsberg, N.; Malmqvist, P.-A. Multiconfiguration Perturbation Theory With Imaginary Level Shift. *Chem. Phys. Lett.* **1997**, *274*, 196–204.
- (13) Finley, J.; Malmqvist, P.-Å.; Roos, B. O.; Serrano-Andrés, L. The Multi-State CASPT2 Method. *Chem. Phys. Lett.* **1998**, *288*, 299–306.
- (14) Shiozaki, T.; Győrffy, W.; Celani, P.; Werner, H.-J. Extended Multi-State Complete Active Space Second-Order Perturbation Theory: Energy and Nuclear Gradients. *J. Chem. Phys.* **2011**, *135*, 081106.
- (15) Battaglia, S.; Lindh, R. Extended Dynamically Weighted CASPT2: The Best of Two Worlds. *J. Chem. Theory Comput.* **2020**, *16*, 1555–1567.
- (16) Andersson, K.; Roos, B. O. Multiconfigurational Second-Order Perturbation Theory: a Test of Geometries and Binding Energies. *Int. J. Quantum Chem.* **1993**, *45*, 591–607.
- (17) Andersson, K. Different Forms of the Zeroth-Order Hamiltonian in Second-Order Perturbation Theory with a Complete Active Space Self-Consistent Field Reference Function. *Theor. Chim. Acta* **1995**, *91*, 31–46.
- (18) Ghigo, G.; Roos, B. O.; Malmqvist, P.-Å. A Modified Definition of the Zeroth-Order Hamiltonian in Multiconfigurational Perturbation Theory (CASPT2). *Chem. Phys. Lett.* **2004**, *396*, 142–149.
- (19) Angeli, C.; Borini, S.; Cavallini, A.; Cestari, M.; Cimraglia, R.; Ferrighi, L.; Sparta, M. Developments in the *N*-Electron Valence State Perturbation Theory. *Int. J. Quantum. Chem.* **2006**, *106*, 686–691.
- (20) Gagliardi, L.; Roos, B. O. Multiconfigurational Quantum Chemical Methods for Molecular Systems Containing Actinides. *Chem. Soc. Rev.* **2007**, *36*, 893–903.
- (21) Siegbahn, P. E. M. Quantum Chemical Studies of Redox-Active Enzymes. *Faraday Discuss.* **2003**, *124*, 289–296.
- (22) Schapiro, I.; Ryazantsev, M. N.; Ding, W. J.; Huntress, M. M.; Melaccio, F.; Andruniow, T.; Olivucci, M. Computational Photobiology and Beyond. *Aust. J. Chem.* **2010**, *63*, 413–429.
- (23) Fouqueau, A.; Casida, M. E.; Daku, L. M. L.; Hauser, A.; Neese, F. Comparison of Density Functionals for Energy And Structural Differences Between the High- [<sup>5</sup>T<sub>2g</sub> : (t<sub>2g</sub>)<sup>4</sup>(e<sub>g</sub>)<sup>2</sup>] And Low- [<sup>1</sup>A<sub>g</sub> : (t<sub>2g</sub>)<sup>6</sup>(e<sub>g</sub>)<sup>0</sup>] Spin States of Iron(II) Coordination Compounds. II. More Functionals and the Hexaminoferrous Cation, [Fe(NH<sub>3</sub>)<sub>6</sub>]<sup>2+</sup>. *J. Chem. Phys.* **2005**, *122*, 044110.
- (24) Pierloot, K.; Vancoillie, S. Relative Energy of the High-(<sup>5</sup>T<sub>2g</sub>) and low-(<sup>1</sup>A<sub>1g</sub>) Spin States of [Fe(H<sub>2</sub>O)<sub>6</sub>]<sup>2+</sup>, [Fe(NH<sub>3</sub>)<sub>6</sub>]<sup>2+</sup>, and [Fe(bpy)<sub>3</sub>]<sup>2+</sup>: CASPT2 versus Density Functional Theory. *J. Chem. Phys.* **2006**, *125*, 124303.
- (25) Neese, F. Theoretical Spectroscopy of Model-Nonheme [Fe(IV)OL<sub>5</sub>]<sup>2+</sup> Complexes in their Lowest Triplet and Quintet States using Multireference ab initio and Density Functional Theory Methods. *J. Inorg. Biochem.* **2006**, *100*, 716 – 726.
- (26) Ordejón, B.; de Graaf, C.; Sousa, C. Light-Induced Excited-State Spin Trapping in Tetrazole-Based Spin Crossover Systems. *J. Am. Chem. Soc.* **2008**, *130*, 13961–13968.
- (27) Suaud, N.; Bonnet, M.-L.; Boilleau, C.; Labèguerie, P.; Guihéry, N. Light-Induced Excited Spin State Trapping: Ab Initio Study of the Physics at the Molecular Level. *J. Am. Chem. Soc.* **2009**, *131*, 715–722.
- (28) Banse, F.; Girerd, J.-J.; Robert, V. Nonheme “FeIVO” Models: Ab Initio Analysis of the Low-Energy Spin State Electronic Structures. *Eur. J. Inorg. Chem.* **2008**, *2008*, 4786–4791.
- (29) Kepenekian, M.; Robert, V.; Le Guennic, B. What Zeroth-Order Hamiltonian for CASPT2 Adiabatic Energetics of Fe(II)N<sub>6</sub> Architectures? *J. Chem. Phys.* **2009**, *131*, 114702.
- (30) Queralt, N.; Taratiel, D.; de Graaf, C.; Caballol, R.; Cimraglia, R.; Angeli, C. On the Applicability of Multireference Second-Order Perturbation Theory to Study Weak Magnetic Coupling in Molecular Complexes. *J. Comput. Chem.* **2008**, *29*, 994–1003.
- (31) Serrano-Andrés, L.; Mechán, M.; Nebot-Gil, I.; Lindh, R.; Roos, B. O. Towards an Accurate Molecular Orbital Theory for Excited States: Ethene, Butadiene, and Hexatriene. *J. Chem. Phys.* **1993**, *98*, 3151–3162.
- (32) Serrano-Andrés, L.; Merchán, M.; Nebot-Gil, I.; Roos, B. O.; Fulscher, M. Theoretical Study of the Electronic Spectra of Cyclopentadiene, Pyrrole, and Furan. *J. Am. Chem. Soc.* **1993**, *115*, 6184–6197.
- (33) Serrano-Andrés, L.; Merchán, M.; Fülischer, M.; Roos, B. O. A Theoretical Study of the Electronic Spectrum of Thiophene. *Chem. Phys. Lett.* **1993**, *211*, 125–134.
- (34) Serrano-Andrés, L.; Merchán, M.; Roos, B. O.; Lindh, R. Theoretical Study of the Internal Charge Transfer in Aminobenzonitriles. *J. Am. Chem. Soc.* **1995**, *117*, 3189–3204.
- (35) Roos, B. O.; Andersson, K.; Fulscher, M. P.; Malmqvist, P.-A.; Serrano-Andrés, L. In *Multiconfigurational Perturbation Theory: Applications in Electronic Spectroscopy*; Prigogine, I., Rice, S. A., Eds.; Wiley, New York, 1996; Vol. 93, pp 219–331.
- (36) Serrano-Andrés, L.; Fülischer, M. P. Theoretical Study of the Electronic Spectroscopy of Peptides. I. The Peptidic Bond: Primary, Secondary, and Tertiary Amides. *J. Am. Chem. Soc.* **1996**, *118*, 12190–12199.
- (37) Serrano-Andrés, L.; Fülischer, M. P.; Roos, B. O.; Merchán, M. Theoretical Study of the Electronic Spectrum of Imidazole. *J. Phys. Chem.* **1996**, *100*, 6484–6491.
- (38) Serrano-Andrés, L.; Fülischer, M. P. Theoretical Study of the Electronic Spectroscopy of Peptides. III. Charge-Transfer Transitions in Polypeptides. *J. Am. Chem. Soc.* **1998**, *120*, 10912–10920.
- (39) Roos, B. O. Theoretical Studies of Electronically Excited States of Molecular Systems Using Multiconfig-

- urational Perturbation Theory. *Acc. Chem. Res.* **1999**, *32*, 137–144.
- (40) Merchan, M.; Serrano-Andrés, L.; Roos, B. O. In *Multiconfigurational Perturbation Theory Applied to Excited States of Organic Compounds*; Recent Advances in Computational Chemistry; World Scientific, 1999; pp 161–195.
- (41) Roos, B. O.; Malmqvist, P.-Å.; Molina, V.; Serrano-Andrés, L.; Merchán, M. Theoretical Characterization of the Lowest-Energy Absorption Band of Pyrrole. *J. Chem. Phys.* **2002**, *116*, 7526–7536.
- (42) Serrano-Andrés, L.; Pou-Amérigo, R.; Fülischer, M. P.; Borin, A. C. Electronic Excited States of Conjugated Cyclic Ketones and Thioketones: A Theoretical Study. *J. Chem. Phys.* **2002**, *117*, 1649–1659.
- (43) Serrano-Andrés, L.; Merchán, M. Quantum Chemistry of the Excited State: 2005 Overview. *J. Mol. Struct. (THEOCHEM)* **2005**, *729*, 99–108.
- (44) Tozer, D. J.; Handy, N. C. Excitation Energies of Benzene from Kohn-Sham Theory. *J. Comput. Chem.* **1999**, *20*, 106–113.
- (45) Burcl, R.; Amos, R. D.; Handy, N. C. Study of Excited States of Furan and Pyrrole by Time-Dependent Density Functional Theory. *Chem. Phys. Lett.* **2002**, *355*, 8–18.
- (46) Peach, M. J. G.; Benfield, P.; Helgaker, T.; Tozer, D. J. Excitation Energies in Density Functional Theory: an Evaluation and a Diagnostic Test. *J. Chem. Phys.* **2008**, *128*, 044118.
- (47) Faber, C.; Boulanger, P.; Duchemin, I.; Attaccalite, C.; Blase, X. Many-Body Greens Function GW and Bethe-Salpeter Study of the Optical Excitations in a Paradigmatic Model Dipeptide. *J. Chem. Phys.* **2013**, *139*, 194308.
- (48) Schreiber, M.; Silva-Junior, M. R.; Sauer, S. P. A.; Thiel, W. Benchmarks for Electronically Excited States: CASPT2, CC2, CCSD and CC3. *J. Chem. Phys.* **2008**, *128*, 134110.
- (49) Silva-Junior, M. R.; Schreiber, M.; Sauer, S. P. A.; Thiel, W. Benchmarks for Electronically Excited States: Time-Dependent Density Functional Theory and Density Functional Theory Based Multireference Configuration Interaction. *J. Chem. Phys.* **2008**, *129*, 104103.
- (50) Sauer, S. P. A.; Schreiber, M.; Silva-Junior, M. R.; Thiel, W. Benchmarks for Electronically Excited States: A Comparison of Noniterative and Iterative Triples Corrections in Linear Response Coupled Cluster Methods: CCSDR(3) versus CC3. *J. Chem. Theory Comput.* **2009**, *5*, 555–564.
- (51) Silva-Junior, M. R.; Thiel, W. Benchmark of Electronically Excited States for Semiempirical Methods: MNDO, AM1, PM3, OM1, OM2, OM3, INDO/S, and INDO/S2. *J. Chem. Theory Comput.* **2010**, *6*, 1546–1564.
- (52) Silva-Junior, M. R.; Sauer, S. P. A.; Schreiber, M.; Thiel, W. Basis Set Effects on Coupled Cluster Benchmarks of Electronically Excited States: CC3, CCSDR(3) and CC2. *Mol. Phys.* **2010**, *108*, 453–465.
- (53) Silva-Junior, M. R.; Schreiber, M.; Sauer, S. P. A.; Thiel, W. Benchmarks of Electronically Excited States: Basis Set Effects on CASPT2 Results. *J. Chem. Phys.* **2010**, *133*, 174318.
- (54) Christiansen, O.; Koch, H.; Jørgensen, P. Response Functions in the CC3 Iterative Triple Excitation Model. *J. Chem. Phys.* **1995**, *103*, 7429–7441.
- (55) Goerigk, L.; Moellmann, J.; Grimme, S. Computation of Accurate Excitation Energies for Large Organic Molecules with Double-Hybrid Density Functionals. *Phys. Chem. Chem. Phys.* **2009**, *11*, 4611–4620.
- (56) Jacquemin, D.; Wathelet, V.; Perpète, E. A.; Adamo, C. Extensive TD-DFT Benchmark: Singlet-Excited States of Organic Molecules. *J. Chem. Theory Comput.* **2009**, *5*, 2420–2435.
- (57) Rohrdanz, M. A.; Martins, K. M.; Herbert, J. M. A Long-Range-Corrected Density Functional That Performs Well for Both Ground-State Properties and Time-Dependent Density Functional Theory Excitation Energies, Including Charge-Transfer Excited States. *J. Chem. Phys.* **2009**, *130*, 054112.
- (58) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C. Assessment of Functionals for TD-DFT Calculations of Singlet-Triplet Transitions. *J. Chem. Theory Comput.* **2010**, *6*, 1532–1537.
- (59) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C.; Valero, R.; Zhao, Y.; Truhlar, D. G. On the Performances of the M06 Family of Density Functionals for Electronic Excitation Energies. *J. Chem. Theory Comput.* **2010**, *6*, 2071–2085.
- (60) Mardirossian, N.; Parkhill, J. A.; Head-Gordon, M. Benchmark Results for Empirical Post-GGA Functionals: Difficult Exchange Problems and Independent Tests. *Phys. Chem. Chem. Phys.* **2011**, *13*, 19325–19337.
- (61) Jacquemin, D.; Perpète, E. A.; Ciofini, I.; Adamo, C. Assessment of the  $\omega$ B97 Family for Excited-State Calculations. *Theor. Chem. Acc.* **2011**, *128*, 127–136.
- (62) Huix-Rotllant, M.; Ipatov, A.; Rubio, A.; Casida, M. E. Assessment of Dressed Time-Dependent Density-Functional Theory for the Low-Lying Valence States of 28 Organic Chromophores. *Chem. Phys.* **2011**, *391*, 120–129.
- (63) Della Sala, F.; Fabiano, E. Accurate Singlet and Triplet Excitation Energies Using the Localized Hartree-Fock Kohn-Sham Potential. *Chem. Phys.* **2011**, *391*, 19–26.
- (64) Trani, F.; Scalmani, G.; Zheng, G. S.; Carnimeo, I.; Frisch, M. J.; Barone, V. Time-Dependent Density Functional Tight Binding: New Formulation and Benchmark of Excited States. *J. Chem. Theory Comput.* **2011**, *7*, 3304–3313.
- (65) Peverati, R.; Truhlar, D. G. Performance of the M11 and M11-L Density Functionals for Calculations of Electronic Excitation Energies by Adiabatic Time-Dependent Density Functional Theory. *Phys. Chem. Chem. Phys.* **2012**, *14*, 11363–11370.
- (66) Domínguez, A.; Aradi, B.; Frauenheim, T.; Lutsker, V.; Niehaus, T. A. Extensions of the Time-

- Dependent Density Functional Based Tight-Binding Approach. *J. Chem. Theory Comput.* **2013**, *9*, 4901–4914.
- (67) Demel, O.; Datta, D.; Nooijen, M. Additional Global Internal Contraction in Variations of Multireference Equation of Motion Coupled Cluster Theory. *J. Chem. Phys.* **2013**, *138*, 134108.
- (68) Schapiro, I.; Sivalingam, K.; Neese, F. Assessment of *n*-Electron Valence State Perturbation Theory for Vertical Excitation Energies. *J. Chem. Theory Comput.* **2013**, *9*, 3567–3580.
- (69) Voityuk, A. A. INDO/X: A New Semiempirical Method for Excited States of Organic and Biological Molecules. *J. Chem. Theory Comput.* **2014**, *10*, 4950–4958.
- (70) Harbach, P. H. P.; Wormit, M.; Dreuw, A. The Third-Order Algebraic Diagrammatic Construction Method (ADC(3)) for the Polarization Propagator for Closed-Shell Molecules: Efficient Implementation and Benchmarking. *J. Chem. Phys.* **2014**, *141*, 064113.
- (71) Yang, Y.; Peng, D.; Lu, J.; Yang, W. Excitation Energies from Particle-Particle Random Phase Approximation: Davidson Algorithm and Benchmark Studies. *J. Chem. Phys.* **2014**, *141*, 124104.
- (72) Sauer, S. P.; Pitzner-Frydendahl, H. F.; Buse, M.; Jensen, H. J. A.; Thiel, W. Performance of SOPPA-Based Methods in the Calculation of Vertical Excitation Energies and Oscillator Strengths. *Mol. Phys.* **2015**, *113*, 2026–2045.
- (73) Piecuch, P.; Hansen, J. A.; Ajala, A. O. Benchmarking the Completely Renormalised Equation-Of-Motion Coupled-Cluster Approaches for Vertical Excitation Energies. *Mol. Phys.* **2015**, *113*, 3085–3127.
- (74) Tajti, A.; Szalay, P. G. Investigation of the Impact of Different Terms in the Second Order Hamiltonian on Excitation Energies of Valence and Rydberg States. *J. Chem. Theory Comput.* **2016**, *12*, 5477–5482.
- (75) Maier, T. M.; Bahmann, H.; Arbuznikov, A. V.; Kaupp, M. Validation of Local Hybrid Functionals for TDDFT Calculations of Electronic Excitation Energies. *J. Chem. Phys.* **2016**, *144*, 074106.
- (76) Rishi, V.; Perera, A.; Nooijen, M.; Bartlett, R. J. Excited States from Modified Coupled Cluster Methods: are they any Better than EOM-CCSD? *J. Chem. Phys.* **2017**, *146*, 144104.
- (77) Dutta, A. K.; Nooijen, M.; Neese, F.; Izsák, R. Exploring the Accuracy of a Low Scaling Similarity Transformed Equation of Motion Method for Vertical Excitation Energies. *J. Chem. Theory Comput.* **2018**, *14*, 72–91.
- (78) Helmich-Paris, B. Benchmarks for Electronically Excited States with CASSCF Methods. *J. Chem. Theory Comput.* **2019**, *15*, 4170–4179.
- (79) Roca-Sanjuan, D.; Aquilante, F.; Lindh, R. Multi-configuration Second-Order Perturbation Theory Approach to Strong Electron Correlation in Chemistry and Photochemistry. *WIREs Comput. Mol. Sci.* **2012**, *2*, 585–603.
- (80) Segarra-Martí, J.; Roca-Sanjuán, D.; Merchán, M.; Lindh, R. On the Photophysics and Photochemistry of the Water Dimer. *J. Chem. Phys.* **2012**, *137*, 244309.
- (81) Li, Q.; Migani, A.; Blancafort, L. Irreversible Phototautomerization of o-Phthalaldehyde Through Electronic Relocation. *Phys. Chem. Chem. Phys.* **2012**, *14*, 6561–6568.
- (82) Gobbo, J. P.; Saurí, V.; Roca-Sanjuán, D.; Serrano-Andrés, L.; Merchán, M.; Borin, A. C. On the Deactivation Mechanisms of Adenine–Thymine Base Pair. *J. Phys. Chem. B* **2012**, *116*, 4089–4097.
- (83) Gobbo, J. P.; Borin, A. C. On the Mechanisms of Triplet Excited State Population in 8-Azaadenine. *J. Phys. Chem. B* **2012**, *116*, 14000–14007.
- (84) Guo, X.; Cao, Z. Low-Lying Electronic States and their Nonradiative Deactivation of Thieno[3,4-b]Pyrazine: an *ab initio* Study. *J. Chem. Phys.* **2012**, *137*, 224313.
- (85) Yamazaki, S.; Taketsugu, T. Nonradiative Deactivation Mechanisms of Uracil, Thymine, and 5-Fluorouracil: A Comparative *ab Initio* Study. *J. Phys. Chem. A* **2012**, *116*, 491–503.
- (86) Gozem, S.; Huntress, M.; Schapiro, I.; Lindh, R.; Granovsky, A. A.; Angeli, C.; Olivucci, M. Dynamic Electron Correlation Effects on the Ground State Potential Energy Surface of a Retinal Chromophore Model. *J. Chem. Theory Comput.* **2012**, *8*, 4069–4080.
- (87) Melaccio, F.; Ferré, N.; Olivucci, M. Quantum Chemical Modeling of Rhodopsin Mutants Displaying Switchable Colors. *Phys. Chem. Chem. Phys.* **2012**, *14*, 12485–12495.
- (88) Gadzhiev, O. B.; Ignatov, S. K.; Krisyuk, B. E.; Maiorov, A. V.; Gangopadhyay, S.; Masunov, A. E. Quantum Chemical Study of the Initial Step of Ozone Addition to the Double Bond of Ethylene. *J. Phys. Chem. A* **2012**, *116*, 10420–10434.
- (89) Gobbo, J. P.; Borin, A. C. On The Population of Triplet Excited States of 6-Aza-2-Thiothymine. *J. Phys. Chem. A* **2013**, *117*, 5589–5596.
- (90) Francés-Monerris, A.; Merchán, M.; Roca-Sanjuán, D. Electronic UV-Vis Transient Spectra of the ·OH Reaction Products of Uracil, Thymine, Cytosine, and 5,6-Dihydrouracil by Using the Complete Active Space Self-Consistent Field Second-Order Perturbation (CASPT2//CASSCF) Theory. *J. Chem. Phys.* **2013**, *139*, 071101.
- (91) Choi, H.; Park, Y. C.; Lee, Y. S.; An, H.; Baeck, K. K. Theoretical Study of the Extremely Small Torsional Barriers of Thiophenol in the Ground and the First Excited Electronic States. *Chem. Phys. Lett.* **2013**, *580*, 32–36.
- (92) Huix-Rotllant, M.; Filatov, M.; Gozem, S.; Schapiro, I.; Olivucci, M.; Ferré, N. Assessment of Density Functional Theory for Describing the Correlation Effects on the Ground and Excited State Potential Energy Surfaces of a Retinal Chromophore Model. *J. Chem. Theory Comput.* **2013**, *9*, 3917–3932.
- (93) Nakayama, A.; Harabuchi, Y.; Yamazaki, S.; Taketsugu, T. Photophysics of Cytosine Tautomers: New

- Insights into the Nonradiative Decay Mechanisms from MS-CASPT2 Potential Energy Calculations and Excited-State Molecular Dynamics Simulations. *Phys. Chem. Chem. Phys.* **2013**, *15*, 12322–12339.
- (94) Komainda, A.; Ostojić, B.; Köppel, H. Ab Initio Quantum Study of Nonadiabatic  $S_1 - S_2$  Photodynamics of *s-trans*-Butadiene. *J. Phys. Chem. A* **2013**, *117*, 8782–8793.
- (95) Perrier, A.; Aloise, S.; Olivucci, M.; Jacquemin, D. Inverse versus Normal Dithienylethenes: Computational Investigation of the Photocyclization Reaction. *J. Phys. Chem. Lett.* **2013**, *4*, 2190–2196.
- (96) Giussani, A. Toward the Understanding of the Photophysics and Photochemistry of 1-Nitronaphthalene under Solar Radiation: The First Theoretical Evidence of a Photodegradation Intramolecular Rearrangement Mechanism Involving the Triplet States. *J. Chem. Theory Comput.* **2014**, *10*, 3987–3995.
- (97) Zobel, J. P.; Nogueira, J. J.; González, L. Quenching of Charge Transfer in Nitrobenzene Induced by Vibrational Motion. *J. Phys. Chem. Lett.* **2015**, *6*, 3006–3011.
- (98) Mai, S.; Marquetand, P.; González, L. A Static Picture of the Relaxation and Intersystem Crossing Mechanisms of Photoexcited 2-Thiouracil. *J. Phys. Chem. A* **2015**, *119*, 9524–9533.
- (99) Nenov, A.; Segarra-Martí, J.; Giussani, A.; Conti, I.; Rivalta, I.; Dumont, E.; Jaiswal, V. K.; Altavilla, S. F.; Mukamel, S.; Garavelli, M. Probing Deactivation Pathways of DNA Nucleobases by Two-Dimensional Electronic Spectroscopy: First Principles Simulations. *Faraday Discuss.* **2015**, *177*, 345–362.
- (100) El-Zohry, A. M.; Roca-Sanjuán, D.; Zietz, B. Ultrafast Twisting of the Indoline Donor Unit Utilized in Solar Cell Dyes: Experimental and Theoretical Studies. *J. Phys. Chem. C* **2015**, *119*, 2249–2259.
- (101) Dumont, E.; Wibowo, M.; Roca-Sanjuán, D.; Garavelli, M.; Assfeld, X.; Monari, A. Resolving the Benzophenone DNA-Photosensitization Mechanism at QM/MM Level. *J. Phys. Chem. Lett.* **2015**, *6*, 576–580.
- (102) Matis, J. R.; Schönborn, J. B.; Saalfrank, P. A Multi-Reference Study Of The Byproduct Formation For A Ring-Closed Dithienylethene Photoswitch. *Phys. Chem. Chem. Phys.* **2015**, *17*, 14088–14095.
- (103) Segado, M.; Gómez, I.; Reguero, M. Intramolecular Charge Transfer in Aminobenzonitriles and Tetrafluoro Counterparts: Fluorescence Explained by Competition Between Low-Lying Excited States and Radiationless Deactivation. Part I: A Mechanistic Overview of the Parent System ABN. *Phys. Chem. Chem. Phys.* **2016**, *18*, 6861–6874.
- (104) Bao, J. J.; Gagliardi, L.; Truhlar, D. G. Multiconfiguration Pair-Density Functional Theory for Doublet Excitation Energies and Excited State Geometries: the Excited States of CN. *Phys. Chem. Chem. Phys.* **2017**, *19*, 30089–30096.
- (105) Giussani, A.; Worth, G. A. Insights into the Complex Photophysics and Photochemistry of the Simplest Nitroaromatic Compound: A CASPT2/CASSCF Study on Nitrobenzene. *J. Chem. Theory Comput.* **2017**, *13*, 2777–2788.
- (106) Heindl, M.; González, L. A XMS-CASPT2 non-Adiabatic Dynamics Study on Pyrrole. *Comput. Theor. Chem.* **2019**, *1155*, 38–46.
- (107) Giussani, A.; Worth, G. A. Similar Chemical Structures, Dissimilar Triplet Quantum Yields: a CASPT2 Model Rationalizing the Trend of Triplet Quantum Yields in Nitroaromatic Systems. *Phys. Chem. Chem. Phys.* **2019**, *21*, 10514–10522.
- (108) Avila-Ferrer, F. J.; Angeli, C.; Cerezo, J.; Coriani, S.; Ferretti, A.; Santoro, F. The Intriguing Case of the One-Photon and Two-Photon Absorption of a Prototypical Symmetric Squaraine: Comparison of TDDFT and Wave-Function Methods. *ChemPhotoChem* **2019**, *3*, 778–793.
- (109) Zobel, J. P.; Nogueira, J. J.; Gonzalez, L. The IPEA Dilemma in CASPT2. *Chem. Sci.* **2017**, *8*, 1482–1499.
- (110) Pierloot, K.; Vancoillie, S. Relative Energy of the High- ( $^5T_{2g}$ ) and Low- ( $^1A_{1g}$ ) Spin states of the Ferrous Complexes [Fe(L)(NHS<sub>4</sub>)]: CASPT2 versus Density Functional Theory. *J. Chem. Phys.* **2008**, *128*, 034104.
- (111) Lawson Daku, L. M.; Aquilante, F.; Robinson, T. W.; Hauser, A. Accurate Spin-State Energetics of Transition Metal Complexes. 1. CCSD(T), CASPT2, and DFT Study of [M(NCH<sub>6</sub>)<sub>6</sub>]<sup>2+</sup> (M = Fe, Co). *J. Chem. Theory Comput.* **2012**, *8*, 4216–4231.
- (112) Rudavskiy, A.; Sousa, C.; de Graaf, C.; Havenith, R. W. A.; Broer, R. Computational Approach to the Study of Thermal Spin Crossover Phenomena. *J. Chem. Phys.* **2014**, *140*, 184318.
- (113) Vela, S.; Fumanal, M.; Ribas-Ariño, J.; Robert, V. On the Zeroth-Order Hamiltonian for CASPT2 Calculations of Spin Crossover Compounds. *J. Comput. Chem.* **2016**, *37*, 947–953.
- (114) Wen, J.; Han, B.; Havlas, Z.; Michl, J. An MS-CASPT2 Calculation of the Excited Electronic States of an Axial Difluoroborondipyrromethene (BODIPY) Dimer. *J. Chem. Theory Comput.* **2018**, *14*, 4291–4297.
- (115) Casida, M. E.; Huix-Rotllant, M. Progress in Time-Dependent Density-Functional Theory. *Annu. Rev. Phys. Chem.* **2012**, *63*, 287–323.
- (116) Ullrich, C. *Time-Dependent Density-Functional Theory: Concepts and Applications*; Oxford Graduate Texts; Oxford University Press: New York, 2012.
- (117) Blase, X.; Duchemin, I.; Jacquemin, D. The Bethe-Salpeter Equation in Chemistry: Relations with TD-DFT, Applications and Challenges. *Chem. Soc. Rev.* **2018**, *47*, 1022–1043.
- (118) Blase, X.; Duchemin, I.; Jacquemin, D.; Loos, P. F. The Bethe-Salpeter Formalism: From Physics to Chemistry. *J. Phys. Chem. Lett.* **2020**, *11*, 7371–7382.
- (119) Dreuw, A.; Wormit, M. The Algebraic Diagrammatic Construction Scheme for the Polarization Propagator for the Calculation of Excited States. *WIREs Comput. Mol. Sci.* **2015**, *5*, 82–95.



- (120) Christiansen, O.; Koch, H.; Jørgensen, P. The Second-Order Approximate Coupled Cluster Singles and Doubles Model CC2. *Chem. Phys. Lett.* **1995**, *243*, 409–418.
- (121) Hättig, C.; Weigend, F. CC2 Excitation Energy Calculations on Large Molecules Using the Resolution of the Identity Approximation. *J. Chem. Phys.* **2000**, *113*, 5154–5161.
- (122) Loos, P. F.; Scemama, A.; Jacquemin, D. The Quest for Highly-Accurate Excitation Energies: A Computational Perspective. *J. Phys. Chem. Lett.* **2020**, *11*, 2374–2383.
- (123) Loos, P.-F.; Scemama, A.; Blondel, A.; Garniron, Y.; Caffarel, M.; Jacquemin, D. A Mountaineering Strategy to Excited States: Highly-Accurate Reference Energies and Benchmarks. *J. Chem. Theory Comput.* **2018**, *14*, 4360–4379.
- (124) Loos, P.-F.; Boggio-Pasqua, M.; Scemama, A.; Caffarel, M.; Jacquemin, D. Reference Energies for Double Excitations. *J. Chem. Theory Comput.* **2019**, *15*, 1939–1956.
- (125) Loos, P.-F.; Lipparini, F.; Boggio-Pasqua, M.; Scemama, A.; Jacquemin, D. A Mountaineering Strategy to Excited States: Highly-Accurate Energies and Benchmarks for Medium Size Molecules. *J. Chem. Theory Comput.* **2020**, *16*, 1711–1741.
- (126) Loos, P.-F.; Jacquemin, D. Is ADC(3) as Accurate as CC3 for Valence and Rydberg Transition Energies? *J. Phys. Chem. Lett.* **2020**, *11*, 974–980.
- (127) Loos, P.-F.; Scemama, A.; Boggio-Pasqua, M.; Jacquemin, D. A Mountaineering Strategy to Excited States: Highly-Accurate Energies and Benchmarks for Exotic Molecules and Radicals. *J. Chem. Theory Comput.* **2020**, *16*, 3720–3736.
- (128) Vèril, M.; Scemama, A.; Caffarel, M.; Lipparini, F.; Boggio-Pasqua, M.; Jacquemin, D.; Loos, P.-F. QUESTDB: a Database of Highly-Accurate Excitation Energies for the Electronic Structure Community. *WIREs Comput. Mol. Sci.* **2021**, *11*, e1517.
- (129) Chrayteh, A.; Blondel, A.; Loos, P.-F.; Jacquemin, D. A Mountaineering Strategy to Excited States: Highly-Accurate Oscillator Strengths and Dipole Moments of Small Molecules. *J. Chem. Theory Comput.* **2021**, *17*, 416–438.
- (130) Loos, P.-F.; Comin, M.; Blase, X.; Jacquemin, D. Reference Energies for Intramolecular Charge-Transfer Excitations. *J. Chem. Theory Comput.* **2021**, *17*, 3666–3686.
- (131) Garniron, Y.; Scemama, A.; Giner, E.; Caffarel, M.; Loos, P.-F. Selected Configuration Interaction Dressed by Perturbation. *J. Chem. Phys.* **2018**, *149*, 064103.
- (132) Scemama, A.; Caffarel, M.; Benali, A.; Jacquemin, D.; Loos, P.-F. Influence of Pseudopotentials on Excitation Energies from Selected Configuration Interaction and Diffusion Monte Carlo. *Res. Chem.* **2019**, *1*, 100002.
- (133) Kucharski, S. A.; Bartlett, R. J. Recursive Intermediate Factorization and Complete Computational Linearization of the Coupled-Cluster Single, Double, Triple, and Quadruple Excitation Equations. *Theor. Chim. Acta* **1991**, *80*, 387–405.
- (134) Kállay, M.; Gauss, J.; Szalay, P. G. Analytic First Derivatives for General Coupled-Cluster and Configuration Interaction Models. *J. Chem. Phys.* **2003**, *119*, 2991–3004.
- (135) Kállay, M.; Gauss, J. Calculation of Excited-State Properties Using General Coupled-Cluster and Configuration-Interaction Models. *J. Chem. Phys.* **2004**, *121*, 9257–9269.
- (136) Hirata, S. Higher-Order Equation-of-Motion Coupled-Cluster Methods. *J. Chem. Phys.* **2004**, *121*, 51–59.
- (137) Noga, J.; Bartlett, R. J. The Full CCSDT Model for Molecular Electronic Structure. *J. Chem. Phys.* **1987**, *86*, 7041–7050.
- (138) Scuseria, G. E.; Schaefer, H. F. A New Implementation of the Full CCSDT Model for Molecular Electronic Structure. *Chem. Phys. Lett.* **1988**, *152*, 382–386.
- (139) Kucharski, S. A.; Włoch, M.; Musiał, M.; Bartlett, R. J. Coupled-Cluster Theory for Excited Electronic States: The Full Equation-Of-Motion Coupled-Cluster Single, Double, and Triple Excitation Method. *J. Chem. Phys.* **2001**, *115*, 8263–8266.
- (140) Kowalski, K.; Piecuch, P. The Active-Space Equation-of-Motion Coupled-Cluster Methods for Excited Electronic States: Full EOMCCSDT. *J. Chem. Phys.* **2001**, *115*, 643–651.
- (141) Kowalski, K.; Piecuch, P. Excited-State Potential Energy Curves of CH<sup>+</sup>: a Comparison of the EOM-CCSDT And Full EOMCCSDT Results. *Chem. Phys. Lett.* **2001**, *347*, 237–246.
- (142) Watts, J. D.; Bartlett, R. J. Iterative and Non-Iterative Triple Excitation Corrections in Coupled-Cluster Methods for Excited Electronic States: the EOM-CCSDT-3 and EOM-CCSD( $\tilde{T}$ ) Methods. *Chem. Phys. Lett.* **1996**, *258*, 581–588.
- (143) Prochnow, E.; Harding, M. E.; Gauss, J. Parallel Calculation of CCSDT and Mk-MRCCSDT Energies. *J. Chem. Theory Comput.* **2010**, *6*, 2339–2347.
- (144) Kendall, R. A.; Dunning, T. H.; Harisson, R. J. Electron Affinities of the First-Row Atoms Revisited. Systematic Basis Sets and Wave Functions. *J. Chem. Phys.* **1992**, *96*, 6796–6806.
- (145) Budzák, Š.; Scalmani, G.; Jacquemin, D. Accurate Excited-State Geometries: a CASPT2 and Coupled-Cluster Reference Database for Small Molecules. *J. Chem. Theory Comput.* **2017**, *13*, 6237–6252.
- (146) Jacquemin, D. What is the Key for Accurate Absorption and Emission Calculations ? Energy or Geometry ? *J. Chem. Theory Comput.* **2018**, *14*, 1534–1543.
- (147) Bremond, E.; Savarese, M.; Adamo, C.; Jacquemin, D. Accuracy of TD-DFT Geometries: a Fresh Look. *J. Chem. Theory Comput.* **2018**, *14*, 3715–3727.
- (148) Werner, H.-J.; Knowles, P. J.; Knizia, G.; Manby, F. R.; Schütz, M. Molpro: a General-Purpose Quantum Chemistry Program Package. *WIREs Comput. Mol. Sci.* **2012**, *2*, 242–253.

- (149) Werner, H.-J. et al. *MOLPRO, Version 2019.2, a Package of ab initio Programs*, 2019, see <https://www.molpro.net>.
- (150) Angeli, C.; Borini, S.; Ferrighi, L.; Cimiraglia, R. Ab Initio *N*-Electron Valence State Perturbation Theory Study of the Adiabatic Transitions in Carbonyl Molecules: Formaldehyde, Acetaldehyde, and Acetone. *J. Chem. Phys.* **2005**, *122*, 114304.
- (151) Kánnár, D.; Tajti, A.; Szalay, P. G. Accuracy of Coupled Cluster Excitation Energies in Diffuse Basis Sets. *J. Chem. Theory Comput.* **2017**, *13*, 202–209.
- (152) Xiao, H.; Maeda, S.; Morokuma, K. CASPT2 Study of Photodissociation Pathways of Ketene. *J. Phys. Chem. A* **2013**, *117*, 7001–7008.
- (153) Send, R.; Valsson, O.; Filippi, C. Electronic Excitations of Simple Cyanine Dyes: Reconciling Density Functional and Wave Function Methods. *J. Chem. Theory Comput.* **2011**, *7*, 444–455.
- (154) Le Guennic, B.; Jacquemin, D. Taking Up the Cyanine Challenge with Quantum Tools. *Acc. Chem. Res.* **2015**, *48*, 530–537.
- (155) Pašteka, L. F.; Melicherčík, M.; Neogrády, P.; Urban, M. CASPT2 and CCSD(T) Calculations of Dipole Moments and Polarizabilities of Acetone in Excited States. *Mol. Phys.* **2012**, *110*, 2219–2237.
- (156) Aquilante, F.; Barone, V.; Roos, B. O. A Theoretical Investigation of Valence and Rydberg Electronic States of Acrolein. *J. Chem. Phys.* **2003**, *119*, 12323–12334.
- (157) Ostojić, B.; Domcke, W. Ab initio Investigation of the Potential Energy Surfaces Involved in the photo-physics of *s-trans*-1,3-Butadiene. *Chem. Phys.* **2001**, *269*, 1–10.
- (158) Saha, B.; Ehara, M.; Nakatsuji, H. Singly and Doubly Excited States of Butadiene, Acrolein, and Glyoxal: Geometries and Electronic Spectra. *J. Chem. Phys.* **2006**, *125*, 014316.
- (159) Watson, M. A.; Chan, G. K.-L. Excited States of Butadiene to Chemical Accuracy: Reconciling Theory and Experiment. *J. Chem. Theory Comput.* **2012**, *8*, 4013–4018.
- (160) Isegawa, M.; Peverati, R.; Truhlar, D. G. Performance of Recent and High-Performance Approximate Density Functionals for Time-Dependent Density Functional Theory Calculations of Valence and Rydberg Electronic Transition Energies. *J. Chem. Phys.* **2012**, *137*, 244104.
- (161) Isegawa, M.; Truhlar, D. G. Valence Excitation Energies of Alkenes, Carbonyl Compounds, and Azabenzenes by Time-Dependent Density Functional Theory: Linear Response of the Ground State Compared to Collinear and Noncollinear Spin-Flip TDDFT with the Tamm-Dancoff Approximation. *J. Chem. Phys.* **2013**, *138*, 134111.
- (162) Schwabe, T.; Goerigk, L. Time-Dependent Double-Hybrid Density Functionals with Spin-Component and Spin-Opposite Scaling. *J. Chem. Theory Comput.* **2017**, *13*, 4307–4323.
- (163) Shu, Y.; Truhlar, D. G. Doubly Excited Character or Static Correlation of the Reference State in the Controversial  $2^1A_g$  State of *trans*-Butadiene? *J. Am. Chem. Soc.* **2017**, *139*, 13770–13778.
- (164) Sokolov, A. Y.; Guo, S.; Ronca, E.; Chan, G. K.-L. Time-Dependent *N*-Electron Valence Perturbation Theory With Matrix Product State Reference Wavefunctions for Large Active Spaces and Basis Sets: Applications to the Chromium Dimer and all-*trans* Polyenes. *J. Chem. Phys.* **2017**, *146*, 244102.
- (165) Chien, A. D.; Holmes, A. A.; Otten, M.; Umrigar, C. J.; Sharma, S.; Zimmerman, P. M. Excited States of Methylene, Polyenes, and Ozone from Heat-Bath Configuration Interaction. *J. Phys. Chem. A* **2018**, *122*, 2714–2722.
- (166) Copan, A. V.; Sokolov, A. Y. Linear-Response Density Cumulant Theory for Excited Electronic States. *J. Chem. Theory Comput.* **2018**, *14*, 4097–4108.
- (167) Tran, T.; Segarra-Martí, J.; Bearpark, M. J.; Robb, M. A. Molecular Vertical Excitation Energies Studied with First-Order RASSCF (RAS[1,1]): Balancing Covalent and Ionic Excited States. *J. Phys. Chem. A* **2019**, *123*, 5223–5230.
- (168) Merchán, M.; González-Luque, R.; Roos, B. O. A Theoretical Determination of the Electronic Spectrum of Methylene-cyclopropane. *Theor. Chim. Acta* **1996**, *94*, 143–154.
- (169) Liu, L.; Xia, S.; Fang, W.-H. Photodecarbonylation Mechanism of Cyclopropenone in the Gas Phase: Electronic Structure Calculation and AIMS Dynamics Simulation. *J. Phys. Chem. A* **2014**, *118*, 8977–8985.
- (170) McDiarmid, R.; Sabljčić, A.; Doering, J. P. Valence Transitions in 1,3-Cyclopentadiene, 1,3-Cyclohexadiene, and 1,3-Cycloheptadiene. *J. Chem. Phys.* **1985**, *83*, 2147–2152.
- (171) Pastore, M.; Angeli, C.; Cimiraglia, R. An Application of Second and Third-Order *n*-Electron Valence State Perturbation Theory to the Calculation of the Vertical Electronic Spectrum of Furan. *Chem. Phys. Lett.* **2006**, *426*, 445–451.
- (172) Li, X.; Paldus, J. A Multireference Coupled-Cluster Study of Electronic Excitations in Furan and Pyrrole. *J. Phys. Chem. A* **2010**, *114*, 8591–8600.
- (173) Holland, D.; Seddon, E.; Trofimov, A.; Gromov, E.; Wormit, M.; Dreuw, A.; Korona, T.; de Oliveira, N.; Archer, L.; Joyeux, D. A Study of the Excited Electronic States of Normal and Fully Deuterated Furan by Photoabsorption Spectroscopy and High-Level ab initio Calculations. *J. Mol. Spectrosc.* **2015**, *315*, 184–195.
- (174) Di Lonardo, G.; Galloni, G.; Trombetti, A.; Zauli, C. Electronic Spectrum of Thiophen and Some Deuterated Thiophens. *J. Chem. Soc., Faraday Trans. 2* **1972**, *68*, 2009–2016.
- (175) Gyorffy, W.; Shiozaki, T.; Knizia, G.; Werner, H.-J. Analytical Energy Gradients for Second-Order Multireference Perturbation Theory Using Density Fitting. *J. Chem. Phys.* **2013**, *138*, 104104.

- (176) Pastore, M.; Angeli, C.; Cimiraglia, R. A Multireference Perturbation Theory Study on the Vertical Electronic Spectrum of Thiophene. *Theor. Chem. Acc.* **2007**, *118*, 35–46.
- (177) Lorentzon, J.; Malmqvist, P.-Å.; Fülischer, M.; Roos, B. O. A CASPT2 Study of the Valence and Lowest Rydberg Electronic States of Benzene and Phenol. *Theor. Chim. Acta* **1995**, *91*, 91–108.
- (178) Sharma, P.; Bernales, V.; Truhlar, D. G.; Gagliardi, L. Valence  $\pi\pi^*$  Excitations in Benzene Studied by Multiconfiguration Pair-Density Functional Theory. *J. Phys. Chem. Lett.* **2019**, *10*, 75–81.
- (179) Doering, J. P. Low-Energy Electron-Impact Study of the First, Second, and Third Triplet States of Benzene. *J. Chem. Phys.* **1969**, *51*, 2866–2870.
- (180) Hiraya, A.; Shobatake, K. Direct Absorption Spectra of Jet-Cooled Benzene in 130–260 nm. *J. Chem. Phys.* **1991**, *94*, 7700–7706.
- (181) Fulscher, M. P.; Andersson, K.; Roos, B. O. Toward an Accurate Molecular Orbital Theory for Excited States: the Azabenzenes. *J. Phys. Chem.* **1992**, *96*, 9204–9212.
- (182) Weber, P.; Reimers, J. R. Ab Initio and Density Functional Calculations of the Energies of the Singlet and Triplet Valence Excited States of Pyrazine. *J. Phys. Chem. A* **1999**, *103*, 9821–9829.
- (183) Woywod, C.; Papp, A.; Halász, G. J.; Vibók, Á. Theoretical Investigation of the Electronic Spectrum of Pyrazine. *Theor. Chem. Acc.* **2010**, *125*, 521–533.
- (184) Sauri, V.; Serrano-Andrés, L.; Moughal Shahi, A. R.; Gagliardi, L.; Vancoillie, S.; Pierloot, K. Multiconfigurational Second-Order Perturbation Theory Restricted Active Space (RASPT2) Method for Electronic Excited States: A Benchmark Study. *J. Chem. Theory Comput.* **2011**, *7*, 153–168.
- (185) Fischer, G.; Wormell, P. Vibronic Analyses of the Lowest Singlet–Singlet and Singlet–Triplet Band Systems of Pyridazine. *Chem. Phys.* **2000**, *257*, 1–20.
- (186) Linert, I.; Zubek, M. Electron Energy-loss Spectroscopy of Excited States of the Diazine Molecules: Pyridazine. *J. Elec. Spect. Rel. Phen.* **2019**, *233*, 69–76.
- (187) Lorentzon, J.; Fülischer, M. P.; Roos, B. O. A Theoretical Study of the Electronic Spectra of Pyridine and Phosphabenzene. *Theor. Chim. Acta* **1995**, *92*, 67–81.
- (188) Walker, I. C.; Palmer, M. H.; Hopkirk, A. The Electronic States of the Azines. II. Pyridine, Studied by VUV Absorption, Near-Threshold Electron Energy Loss Spectroscopy and Ab Initio Multi-Reference Configuration Interaction Calculations. *Chem. Phys.* **1990**, *141*, 365–378.
- (189) Linert, I.; Zubek, M. Electron Energy-Loss Spectroscopy of Excited States of the Pyridine Molecules. *Eur. J. Phys. D* **2016**, *70*, 74.
- (190) Serrano-Andrés, L.; Fülischer, M. P.; Karlström, G. Solvent Effects on Electronic Spectra Studied by Multiconfigurational Perturbation Theory. *Int. J. Quantum Chem.* **1997**, *65*, 167–181.
- (191) Fischer, G.; Cai, Z.-L.; Reimers, J. R.; Wormell, P. Singlet and Triplet Valence Excited States of Pyrimidine. *J. Phys. Chem. A* **2003**, *107*, 3093–3106.
- (192) Rubio, M.; Roos, B. O. A Theoretical Study of the Electronic Spectrum of *s*-Tetrazine. *Mol. Phys.* **1999**, *96*, 603–615.
- (193) Devarajan, A.; Gaenko, A. V.; Khait, Y. G.; Hoffmann, M. R. Generalized van Vleck Perturbation Theory (GVVPT2) Study of the Excited States of Benzene and the Azabenzenes. *J. Phys. Chem. A* **2008**, *112*, 2677–2682.
- (194) Angeli, C.; Cimiraglia, R.; Cestarrì, M. A Multireference *n*-electron Valence State Perturbation Theory Study of the Electronic Spectrum of *s*-tetrazine. *Theor. Chem. Acc.* **2009**, *123*, 287–298.
- (195) Oliva, J. M.; Azenha, M. E. D. G.; Burrows, H. D.; Coimbra, R.; Seixas de Melo, J. S.; Canle L., M.; Fernández, M. I.; Santaballa, J. A.; Serrano-Andrés, L. On the Low-Lying Excited States of sym-Triazine-Based Herbicides. *ChemPhysChem* **2005**, *6*, 306–314.
- (196) Hättig, C. Structure Optimizations for Excited States with Correlated Second-Order Methods: CC2 and ADC(2). In *Response Theory and Molecular Properties (A Tribute to Jan Lindenberg and Poul Jørgensen)*; Jensen, H. A., Ed.; Academic Press, 2005; Vol. 50, pp 37–60.
- (197) Watson, T. J.; Lotrich, V. F.; Szalay, P. G.; Perera, A.; Bartlett, R. J. Benchmarking for Perturbative Triple-Excitations in EE-EOM-CC Methods. *J. Phys. Chem. A* **2013**, *117*, 2569–2579.
- (198) Kánnár, D.; Szalay, P. G. Benchmarking Coupled Cluster Methods on Valence Singlet Excited States. *J. Chem. Theory Comput.* **2014**, *10*, 3757–3765.

## Graphical TOC Entry

