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► **To cite this version:**

Yiying Zhang, Hui Li, Jianhong Chang, Xinyu Guan, Lingxue Q Tang, et al.. 3D Thioether-Based Covalent Organic Frameworks for Selective and Efficient Mercury Removal. *Small*, 2021, 17 (22), pp.2006112. 10.1002/sml.202006112 . hal-03426048

HAL Id: hal-03426048

<https://hal.science/hal-03426048>

Submitted on 11 Nov 2021

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Three-Dimensional Thioether-Based Covalent Organic Frameworks for Selective and Efficient Mercury Removal

Yiying Zhang, Hui Li, Jianhong Chang, Xinyu Guan, Lingxue Tang, Qianrong Fang,
Valentin Valtchev, Yushan Yan, and Shilun Qiu**

Y. Zhang, H. Li, J. Chang, X. Guan, L. Tang, Prof. Q. Fang, Prof. S. Qiu,
State Key Laboratory of Inorganic Synthesis and Preparative Chemistry, Jilin
University, Changchun 130012, P. R. China.

E-mail: qrfang@jlu.edu.cn or sqiu@jlu.edu.cn

Prof. V. Valtchev

Qingdao Institute of Bioenergy and Bioprocess Technology, Chinese Academy of
Sciences, 189 Songling Road, Laoshan District, Qingdao, Shandong 266101, P. R.
China

Prof. V. Valtchev

Normandie Univ, ENSICAEN, UNICAEN, CNRS, Laboratoire Catalyse et
Spectrochimie, 6 Marechal Juin, 14050 Caen, France

Prof. Y. Yan

Department of Chemical and Biomolecular Engineering, Center for Catalytic Science
and Technology, University of Delaware, Newark, DE 19716, USA

Abstract:

Developing functionalized 3D covalent organic frameworks (3D COFs) is critical to broaden their potential applications. However, the introduction of specific functionality in 3D COFs remains a great challenge because most of the functional groups are not compatible with the synthesis conditions. Herein, for the first time 3D thioether-based COFs (JUC-570 and JUC-571) for mercury (Hg^{2+}) removal from aqueous solution is reported. These 3D thioether-based COFs prepared by the bottom-up approach display high Hg^{2+} uptakes (972 mg g^{-1} for JUC-570 and 970 mg g^{-1} for JUC-571 at $\text{pH} = 5$), fast adsorption kinetics (distribution coefficient K_d value of $2.29 \times 10^7 \text{ mL g}^{-1}$ for JUC-570 and $2.07 \times 10^7 \text{ mL g}^{-1}$ for JUC-571), and favorable selectivity. In particular, JUC-570 is periodically decorated with isopropyl groups around imine bonds that markedly improve its chemical stability and effectively prevent the pore collapse, and thus endows high Hg^{2+} adsorption capacity (619 mg g^{-1}) and excellent cycle performance even at $\text{pH} = 1$. This study not only puts forward a new route to construct stable functionalized 3D COFs, but also promotes their potential applications in areas related to the environment.

Keywords: covalent organic framework, mercury removal, selective adsorption, thioether functionalization

1. Introduction

Covalent organic frameworks (COFs) are a type of burgeoning crystalline porous polymers constructed by organic building blocks regularly linked via strong covalent bonds.^[1] Their high porosity, large surface areas, excellent chemical/thermal stabilities, and structurally adjustable reticular skeletons are the basis of different applications,^[2] ranging from gas adsorption and separation,^[3] catalysis,^[4] sensing,^[5] optoelectronics,^[6] drug delivery,^[7] and many others.^[8] The introduction of specific functional groups in COF materials allows tailoring their interactions with guest species for the target application. To date, however, most of the applications were focused on the two-dimensional (2D) COFs with layered eclipsed stacking structures. By contrast, the exploration of three-dimensional (3D) COFs, is still extremely limited.^[9] This is mainly due to the related synthetic challenges,^[10] structural instability,^[11] narrow microporous channels,^[12] and complicated structure determination.^[13] In spite of these issues, 3D COFs have been considered as a powerful design platform for future applications.^[14] We have recently reported some 3D functionalized COFs by elaborate pre-design, such as 3D ionic COFs for selective ion exchange,^[14] 3D Salphen-based COFs as catalytic antioxidants,^[15] 3D tetrathiafulvalene-based COFs for tunable electrical conductivity,^[16] and so on.^[17] It must be noted that the functionalization of 3D COFs still remains largely unexplored. Thus, the development of new 3D functionalized COF materials is vital to enrich the structural diversity and expand their potential applications.

Mercury (Hg) is one of the most harmful heavy metal due to its high toxicity, bioaccumulation, and persistence.^[18] Particularly, excessive Hg species in the environment will cause severe damages to human beings and other living organisms. Most Hg species exist in water in the form of Hg^{2+} , which can be easily converted to more toxic organic mercury species through biological activities.^[19] Consequently, the elimination of the Hg^{2+} from the water-body is a long-term global subject and an urgent stupendous challenge. Owing to the diversity of skeleton structure and pore chemical environment, COFs can be designed as functional target materials that are promising for handling environmental issues. Generally, there are two ways to achieve this goal, namely either the bottom-up approach^[20] or post-synthetic modification.^[21] However, the latter method has nonnegligible shortcomings, which including complicated procedure,^[22] structural breaking,^[23] and uncontrollable quantity of functional site.^[24] On the contrary, the above issues can be effectively avoided through the bottom-up approach that means the direct synthesis of COFs from pre-designed precursors with specific functional groups. At present, only some 2D COFs have been used for Hg^{2+} removal.^[25] Compared to 2D COFs with one-dimensional (1D) channels, 3D COFs not only feature more opening structures, but also possess numerous accessible, functional sites and higher surface areas, which are highly beneficial in Hg^{2+} adsorption. However, there is no related work has been reported yet.

Taking these issues into account, herein, we report 3D functionalized COFs (JUC-570 and JUC-571, JUC = Jilin University China), constructed from thioether

functional units, 3,3',5,5'-tetraisopropyl-2',5'-bis((methylthio)methyl)-[1,1':4',1''-terphenyl]-4,4''-diamine (TBTD) or 2',5'-bis((methylthio)methyl)-[1,1':4',1''-terphenyl]-4,4''-dicarbaldehyde (BTD) by bottom-up approach. More importantly, JUC-570 was fully incorporated with isopropyl groups, which can effectively improve the chemical stability and avoid the pore shrinkage. Both COFs demonstrate high uptakes (972 mg g^{-1} for JUC-570 and 970 mg g^{-1} for JUC-571 at $\text{pH} = 5$), fast adsorption kinetics (distribution coefficient K_d value of $2.29 \times 10^7 \text{ mL g}^{-1}$ for JUC-570 and $2.07 \times 10^7 \text{ mL g}^{-1}$ for JUC-571), and excellent selectivity for Hg^{2+} ion. In particular, JUC-570 decorated with alkyl groups displays high Hg^{2+} adsorption capacity (619 mg g^{-1}), and can be recycled at least four times at $\text{pH} = 1$. To the best of our knowledge, this study represents the first case of 3D thioether-based COFs by the bottom-up approach and their use for the removal of toxic heavy metal ion (Hg^{2+}).

2. Results and Discussion

2.1 Structural Design

Our strategy is based on the flexible thioether groups with small size, which likely possess the compatibility with the synthesis conditions of 3D COFs, and isopropyl groups around imine bonds in the framework that can increase the chemical stability of the material. As shown in **Scheme 1**, the tetrakis(4-formylphenyl)silane (TFS, Scheme 1a) and tetrakis(4-aminophenyl)methane (TAM, Scheme 1b) were designed as tetrahedral knots, and TBTD with isopropyl groups (Scheme 1c) or BTD without isopropyl groups (Scheme 1d) were chosen as two kinds of linear linkers. The

condensation of TFS and TBTD or TAM and BTD resulted in JUC-570 (Scheme 1e) and JUC-571 (Scheme 1f). Linking *C2* and *Td* symmetrical building units results in 3D networks with the **dia** topology (Scheme 1g).^[26] Owing to the tetrahedral centers separated by long linear linkers, the resulting structures tend to be interpenetrated networks.^[27]

2.2. Synthesis and Characterization

3D thioether-based COFs were synthesized from TFS and TBTD or TAM and BTD under solvothermal reaction conditions in a mixture of 1,4-dioxane and mesitylene or a pure 1,4-dioxane in the presence of 6 M acetic acid followed by heating at 120 °C for 3 days. Multiple characterizations were subsequently employed for their detailed structural definition. The morphology of both COFs was examined by scanning electron microscopy (SEM) (Figure S1) and transmission electron microscopy (TEM) (Figure S2), which revealed homogeneous ball shape for JUC-570 and rice shape for JUC-571, respectively. Peaks around 1643 cm^{-1} for JUC-570 and 1624 cm^{-1} for JUC-571 in Fourier transform infrared (FT-IR) spectra indicated the formation of C=N bond. The concomitant disappearance of the N-H stretching vibration of TBTD (3490~3407 cm^{-1}) and TAM (3395~3174 cm^{-1}), as well as the C=O stretching vibration of TFS (1706 cm^{-1}) and BTD (1696 cm^{-1}) gave a piece of direct evidence for the formation of imine linkages (Figures S3 and S4). Solid-state ^{13}C cross-polarization/magic-angle-spinning (CP/MAS) NMR spectroscopy further verified the presence of imine bonds by the peak at 162 ppm for JUC-570 or 159 ppm

for JUC-571 (Figures S5 and S6). Thermogravimetric analysis (TGA) revealed that both COFs were stable up to about 350 °C under N₂ atmosphere (Figures S7 and S8).

The crystallinity and unit cell parameters of 3D thioether-based COFs were determined by powder X-ray diffraction (PXRD) measurements and coupled with the structural simulation using Materials Studio software package (**Figure 1**).^[28] After the geometrical energy minimization, both COFs were set up based on a 10-fold interpenetrated **dia** net. Subsequently their unit cell parameters were acquired ($a = b = 30.2406 \text{ \AA}$, $c = 7.0039 \text{ \AA}$ and $\alpha = \beta = \gamma = 90^\circ$ for JUC-570; $a = b = 28.8381 \text{ \AA}$, $c = 7.1132 \text{ \AA}$ and $\alpha = \beta = \gamma = 90^\circ$ for JUC-571). Besides, the Pawley refinement yielded a PXRD pattern in which peaks at 4.38, 6.19, 6.92, and 9.30° for JUC-570 correspond to the (110), (200), (210), and (300) Bragg peaks of the space group *P*-4 (No. 81). Meanwhile, peaks at 4.41, 6.25, 7.01, 8.87, 12.62, 17.86, and 18.96° for JUC-571 belong to the (110), (200), (210), (300), (400), (440), and (600) Bragg peaks of space group *P*-4 (No. 81). These results were in good agreement with the experimentally observed patterns, indicating by their negligible difference ($R_{wp} = 2.90 \%$ and $R_p = 2.08 \%$ for JUC-570; $R_{wp} = 2.59 \%$ and $R_p = 1.44 \%$ for JUC-571). It should be noted that a similar structure with 10-fold interpenetrated **dia** net (LZU-79) has been proved by a single crystal XRD analysis, and the PXRD patterns of JUC-570 and JUC-571 were well consistent with that of LZU-79 (Figures S9 and S10).^[29] On the basis of the above results, JUC-570 and JUC-571 were proposed to have the expected 10-fold interpenetrated **dia** topology, and the overall frameworks had 1D rectangular channels (**Figure 2**). The formation of highly interpenetrated structures is an expected

consequence of the relatively long linear linkers employed. The interpenetration of the linkers provides the low energy state necessary for the structural stability and generate huge diamond cavities.^[17c]

The nitrogen (N₂) adsorption-desorption isotherms were conducted at 77 K to evaluate the porosity of 3D thioether-based COFs (**Figure 3**). Both materials showed a sharp uptake at a low pressure of $P/P_0 < 0.05$, which fits the typical type I sorption model. The inclination of isotherms in the 0.8-1.0 P/P_0 range and small hysteresis can be attributed to the presence of textural mesopores, which is a consequence of the agglomeration of COF crystals.^[7b] The Brunauer-Emmett-Teller (BET) equation was carried out in the $0.05 < P/P_0 < 0.30$ range, indicating BET specific surface area of 1227 m² g⁻¹ for JUC-570. Pore size distributions calculated by nonlocal density functional theory (NLDFT) showed micropores with a dominant size of 1.09 nm for JUC-570 (Figure 3a inset and Figure S11), which is in good agreement with that predicted from its crystal structure (1.17 nm for JUC-570). Obviously, under the support of isopropyl groups, JUC-570 effectively avoids the common breathing effect phenomenon in 3D COFs.^[30] By contrast, JUC-571 suffered from pore shrinkage after the solvent guests removal owing to the lack of isopropyl groups, which leads to lower BET specific surface area (254 m² g⁻¹) and pore size (0.98 nm, Figure 3b inset and Figure S12).^[31]

The chemical stabilities of 3D thioether-based COFs were also studied by exposure to different environments. As shown in **Figure 4**, JUC-570 still exhibited intense PXRD patterns after treated with 2 M HCl, 2 M HNO₃, 15 M NaOH, boiling water,

and common organic solvents for one week (Figure 4a and Figure S13), which indicates that its crystalline structure is well retained under these conditions. However, JUC-571 was not stable using diluted acid and base as 0.1 M HCl, 0.1 M HNO₃, 1 M NaOH, and boiling water for 24 h, revealed by the PXRD analysis before and after the treatment (Figure 4b and Figure S14). Notably, the difference in chemical stability between two COFs can be attributed to the presence of isopropyl groups in the framework of JUC-571, which can effectively protect the hydrolytically susceptible backbones.^[30, 32]

2.3. Mercury Sorption Studies

2.3.1 Influence of pH

Giving the high porosity, excellent stability, abundant and accessible thioether arms of 3D thioether-based COFs, we conducted a sequence of systematic adsorption experiments to examine their adsorption ability of Hg²⁺ (**Figure 5**). The pH plays a key role in the adsorption process and affects the capture capacity of the adsorbents.^[33] Therefore, we first explored the effect of different pH values on the Hg²⁺ removal of 3D COFs. Because insoluble Hg(OH)₂ can be formed at pH > 6, we explored the pH range of 1-6 for JUC-570. The pH range was limited between 2 and 6 for JUC-571 due to its unstable structure at pH = 1. It was found out that the optimal pH value for Hg²⁺ adsorption on 3D thioether-based COFs is 5 (Figure 5a). Both COFs showed reduced Hg²⁺ capture capacity at other pH values, since the adsorption process is related to the nature of Hg²⁺ species formed at different pH values. At lower pH, the protonation effect on the surface of adsorbent, which produces electrostatic

impediment repulsion with Hg^{2+} and superfluous H^+ , offers sorption competition with Hg^{2+} , and thus prohibits further adsorption.^[34] At higher pH conditions, however, Hg^{2+} begins to undergo hydrolysis, reducing the adsorption amount of Hg^{2+} .^[35] Accordingly, a pH = 5 was considered as an optimum condition for subsequent experiments. It is worth noting that even at very low pH (pH = 1), the Hg^{2+} adsorption capacity of JUC-570 is still as high as 619 mg g^{-1} .

2.3.2 Adsorption Isotherms

To evaluate the Hg^{2+} adsorption capacity of 3D thioether-based COFs, Hg^{2+} solutions with the initial concentrations in the range of $25 \sim 1000 \text{ mg L}^{-1}$ were employed. After the experiment, the COF material was separated and the Hg concentration in the solution analyzed by ICP. As illustrated in Figure 5b, the adsorption capacity of both COFs increased significantly with the increase of the Hg^{2+} ion concentration driven by the concentration gradient force. The experimental maximum Hg^{2+} adsorption capacity of JUC-570 and JUC-571 was calculated to be 972 mg g^{-1} and 970 mg g^{-1} , respectively, which ranks as one of the top values in most typical benchmark adsorbents, such as porous carbon (518 mg g^{-1}),^[36] TAPB-BMTTPA-COF (734 mg g^{-1}),^[25b] Bio-MOF (900 mg g^{-1}),^[37] and PAF-1-SH (1014 mg g^{-1}).^[38] Both equilibrium adsorption isotherms were well-fitted with the Langmuir model that yields a correlation coefficient of 0.998 for JUC-570 (Figure S15) and 0.996 for JUC-571 (Figure S16). We reasoned that the outstanding capacity of JUC-570 stems from the synergistic effect of tremendous affinity between sulfur sites and Hg^{2+} together with high density and accessibility of thioether groups in inherent opening

3D channels. In addition, their PXRD patterns revealed that both materials remained crystalline and structurally intact after adsorbing Hg^{2+} (Figures S17 and S18).

2.3.3 Adsorption Kinetics

To further reveal adsorption process, the adsorption kinetics of Hg^{2+} uptake was investigated in a system with 100.0 mL of $10.0 \text{ mg L}^{-1} \text{ Hg}^{2+}$ and 5.0 mg of 3D thioether-based COFs by measuring the Hg^{2+} concentration at different time. As shown in Figure 5c, JUC-570 can attain 99% of the adsorption capacity at equilibrium within 20 min. The fast adsorption kinetics of JUC-570 can be attributed to flexible short thioether chains, which is beneficial to expose sulfur sites and endows the channel of COFs maximum space.^[22b] In comparison, the adsorption process of JUC-571 is relatively slow. This may be due to the breathing effect of JUC-571, and related slow rate of pore opening in the water. Both Hg^{2+} adsorption data of 3D thioether-based COFs were well fitted with pseudo-second-order kinetic model, and the correlation coefficient as high as 0.9999 for JUC-570 and 0.9996 for JUC-571 (Figures S19 and S20). Distribution coefficient (K_d) represents an important index to evaluate the sorbent affinity to a metal ion,^[25b] was calculated by the equation:

$$K_d = \frac{C_0 - C_e}{C_e} \times \frac{V}{m} \quad (1)$$

Where V is the volume of the treated solution (mL), m is the mass of adsorbent (g), C_0 and C_e are the initial and equilibrium concentration of Hg^{2+} , respectively. Impressively, the K_d values of 3D thioether-based COFs were calculated to equal $2.29 \times 10^7 \text{ mL g}^{-1}$ for JUC-570 and $2.07 \times 10^7 \text{ mL g}^{-1}$ for JUC-571, which can compare to most typical benchmark adsorbents for Hg^{2+} capture, including commercial resins (5.1

$\times 10^5 \text{ mL g}^{-1}$),^[39] TAPB-BMTTPA-COF ($7.82 \times 10^5 \text{ mL g}^{-1}$),^[25b] Thiol-HKUST-1 ($4.73 \times 10^5 \text{ mL g}^{-1}$),^[40] and PAF-1-SH ($5.76 \times 10^7 \text{ mL g}^{-1}$).^[38]

2.3.4 Selectivity

Selectivity tests were performed in a mixed solution containing Hg^{2+} , Cu^{2+} , Ca^{2+} , Mg^{2+} , Fe^{2+} , Zn^{2+} , Ni^{2+} and K^+ for each concentration of 10.0 mg L^{-1} at $\text{pH} = 5$. As described in Figure 5d, both 3D thioether-based COFs could effectively remove Hg^{2+} and exhibited negligible capture capability for various nontoxic competitive ions. Therefore, the interference study indicated that 3D thioether-based COFs can remove Hg^{2+} from aqueous solutions selectively and effectively. The high selectivity is due to a strong affinity between sulfur and mercury since the equilibrium constant between the sulfur atom and mercury is significantly higher than between the sulfur other metals.^[41]

2.3.5 Cycle Performance

Cycle performance of the adsorbent is a critical index for cutting down the cost of the adsorption process in a practical application. 3D thioether-based COFs can be regenerated by treating with 1,2-ethanedithiol and used for cycle tests. As shown in Figure 5e, due to the high chemical stability, JUC-570 exhibited a similar Hg^{2+} removal efficiency after four successive experiments at $\text{pH} = 5$. However, under the same conditions, JUC-571 showed a gradually decreasing Hg^{2+} removal efficiency upon cycling test. It is worth mentioning that, even under the condition of $\text{pH} = 1$, JUC-570 can still retain 94% removal efficiency after four cycles (Figure 5f). In

addition, the crystallinity of reused 3D thioether-based COFs were negligibly declined by PXRD analysis (Figure S21 and Figure S22). These results demonstrated that JUC-570 is robust enough, which makes it a promising candidate in practical water remediation.

2.3.6 Adsorption Mechanism

To shed light on the adsorption mechanism staying behind the outstanding performance of 3D thioether-based COFs in mercury capture, we utilized ^{13}C CP/MAS NMR to figure out the interaction between Hg^{2+} and frameworks. The signals at 14.5 ppm and 36.2 ppm for JUC-570 (Figure S23, purple curve) and at 15.3 ppm and 35.7 ppm for JUC-571 (Figure S24, green curve) were ascribed to the methyl carbon and methylene carbon adjacent to the S atoms. Upon the capture of Hg^{2+} , the carbon signals were shifted to 18.4 ppm and 39.3 ppm for JUC-570 (Figure S23, pink curve) as well as 19.4 ppm and 38.9 ppm for JUC-571 (Figure S24, pink curve), respectively. At the same time, the imine-bond carbons and other signals were almost unchanged. These results pointed out that the flexible thioether groups interact with Hg^{2+} in both 3D thioether-based COFs. In contrast, the imine bonds played a limited role in the Hg^{2+} capture.

3. Conclusion

In conclusion, we rationally designed and synthesized two 3D thioether-based COFs JUC-570 and JUC-571, with 10-fold interpenetrated **dia** topology by the bottom-up approach. By introducing isopropyl groups around imine bonds, JUC-570 showed

much higher chemical stability and BET specific surface area ($1227 \text{ m}^2 \text{ g}^{-1}$). Owing to the thioether functional groups in their frameworks, both COFs had high uptakes (972 mg g^{-1} for JUC-570 and 970 mg g^{-1} for JUC-571 at $\text{pH} = 5$), high adsorption kinetics ($2.29 \times 10^7 \text{ mL g}^{-1}$ for JUC-570 and $2.07 \times 10^7 \text{ mL g}^{-1}$ for JUC-571), and outstanding selectivity for Hg^+ ion. More important, JUC-570 displayed a high Hg^+ adsorption capacity (619 mg g^{-1}), and can be recycled at least four times under highly acidic conditions ($\text{pH} = 1$) due to its excellent stability. This study not only established an efficient route to prepare stable 3D functionalized COFs, but also explored their application in the field of toxic metal decontamination.

4. Experimental Section

Synthesis of JUC-570: A Pyrex tube (volume: ca. 20.0 ml with a body length of 18.0 cm) was charged with TBTD (21.76 mg, 0.04 mmol) and TFS (8.96 mg, 0.02 mmol) in a mixed solution of dioxane (0.5 mL), mesitylene (0.5 mL), and acetic acid (0.2 mL, 6 M). The tube was flash-frozen at 77 K (LN_2 bath), evacuated to an internal pressure of 0.15 mmHg, and flame sealed. Upon sealing, the length of the tube was reduced to ca. 13 cm. Upon warming to room temperature, the tube was heated at $120 \text{ }^\circ\text{C}$ for 3 days to afford a yellow precipitate, which was isolated by filtration and washed with acetone ($3 \times 5.0 \text{ ml}$) and yield of 75% for JUC-570. Analytically calculated for $\text{C}_{96}\text{H}_{112}\text{N}_4\text{S}_4\text{Si}$: C: 96.00; H: 7.64; N: 3.79; S: 8.67; Si: 1.90. Found: C: 95.82; H: 7.69; N: 3.82; S: 8.69; Si: 1.85.

Synthesis of JUC-571: In a manner similar to the preparation of JUC-570, a Pyrex tube (volume: ca. 20.0 mL with a body length of 18.0 cm) was charged with BTB

(16.24 mg, 0.04 mmol) and TAM (7.6 mg, 0.02 mmol) in a mixed solution of dioxane (1.0 mL) and acetic acid (0.2 mL, 6 M). The tube was flash-frozen at 77 K (LN₂ bath), evacuated to an internal pressure of 0.15 mmHg, and flame sealed. Upon sealing, the length of the tube was reduced to ca. 13 cm. The reaction mixture was heated at 120 °C for 3 days to afford a yellow precipitate which was isolated by filtration and washed with acetone (3 × 5.0 ml) and yield of 82% for JUC-571. Annalistically calculated for C₇₃H₆₄N₄S₄: C: 77.90; H: 5.73; N: 4.98; S: 11.39. Found: C: 77.95; H: 5.76; N: 4.89; S: 11.32.

Supporting Information

Supporting Information is available from the Wiley Online Library or from the author.

Acknowledgements

This work was supported by National Natural Science Foundation of China (21621001, 21390394, and 21673094), "111" project (BP0719036 and B17020), and the program for JLU Science and Technology Innovative Research Team, and the Computing Center of Jilin Province for supercomputer time. V.V., S.Q. and Q.F. acknowledge funding from the French-Sino International Laboratory (LIA) "Zeolites".

Conflict of Interest

The authors declare no conflict of interest.

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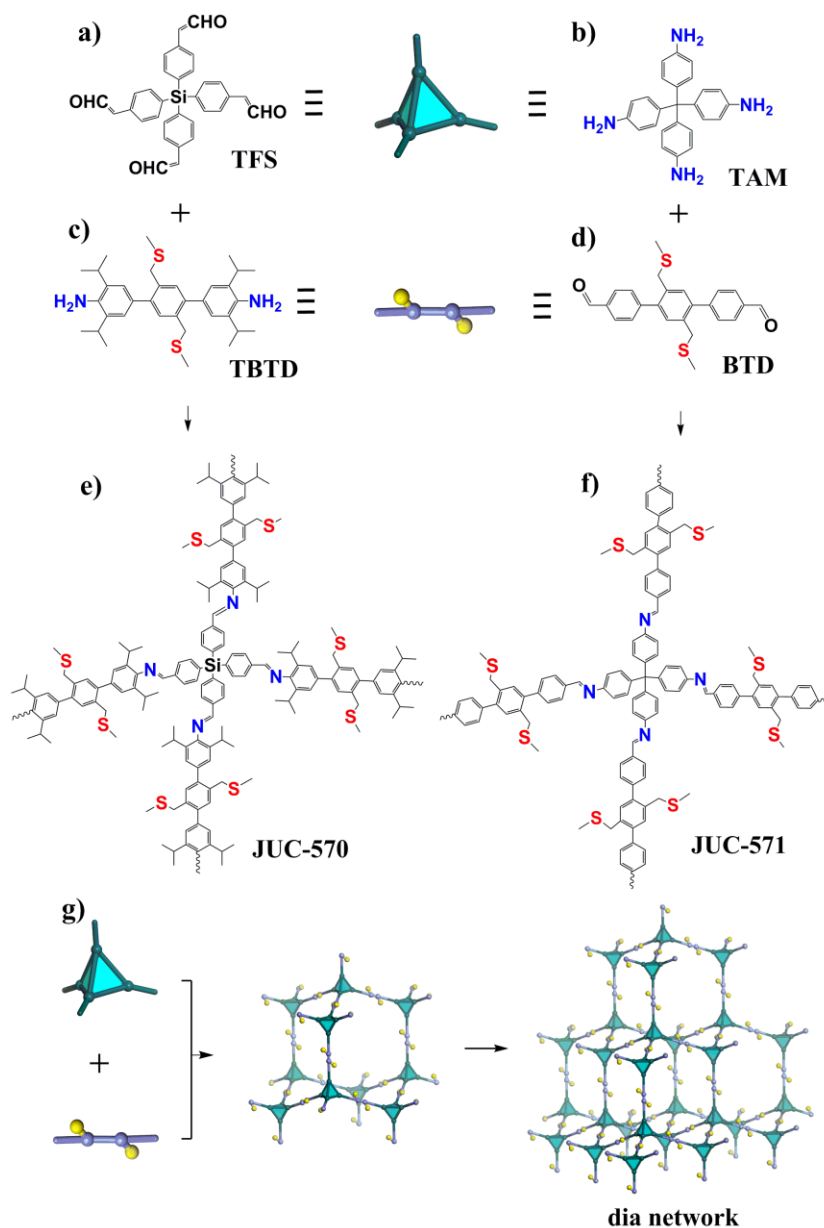
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Figures and captions:



Scheme 1. Scheme of the strategy for preparing 3D thioether-based COFs. Molecular structures of TFS (a) and TAM (b) as tetrahedral building units, TBTD (c) and BTD (d) as linear building units. 3D thioether-based COFs denoted as JUC-570 (e) and JUC-571 (f), are constructed by the condensation reaction of TFS and TBTD, TAM and BTD.

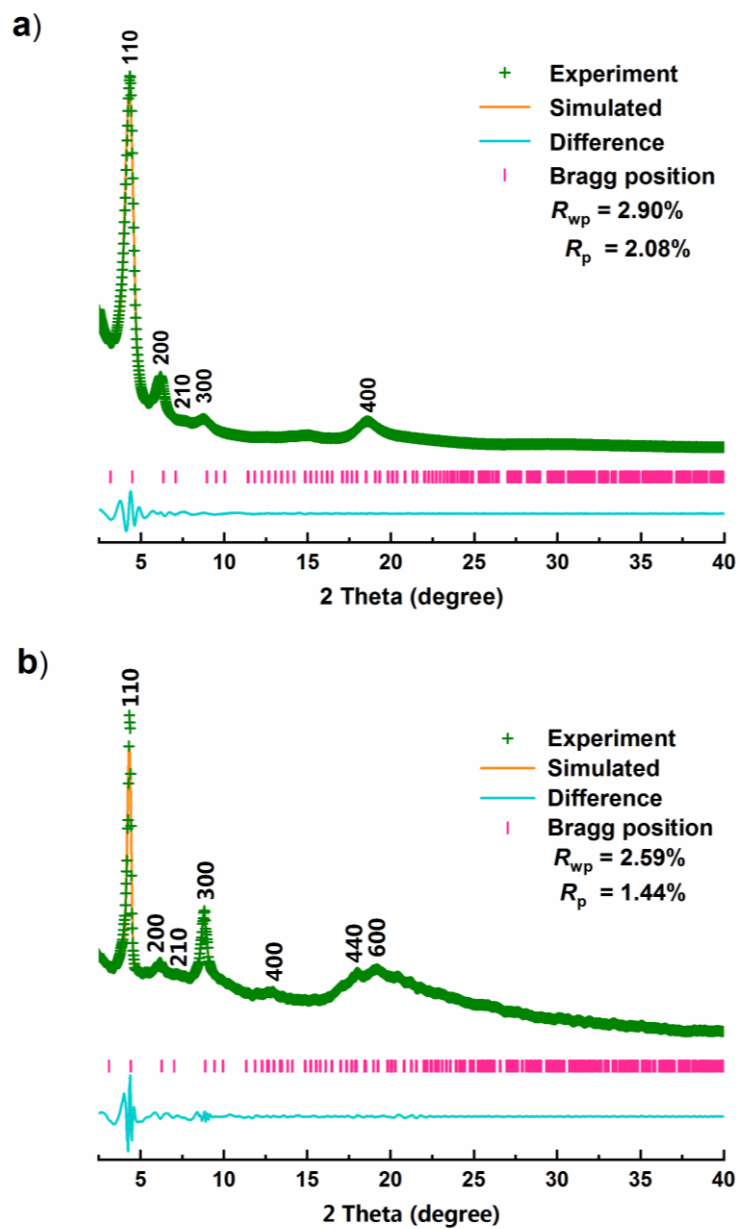


Figure 1. Experimental and refined PXRD patterns of JUC-570 (a) and JUC-571 (b).

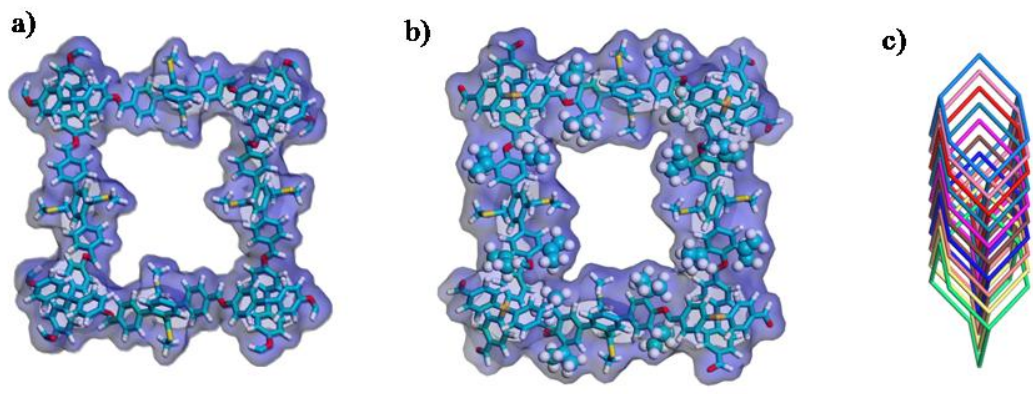


Figure 2. Extended structures of JUC-570 (a) and JUC-571 decorated with isopropyl groups (b). The 10-fold interpenetrated **dia** topology in JUC-570 or JUC-571 (c).

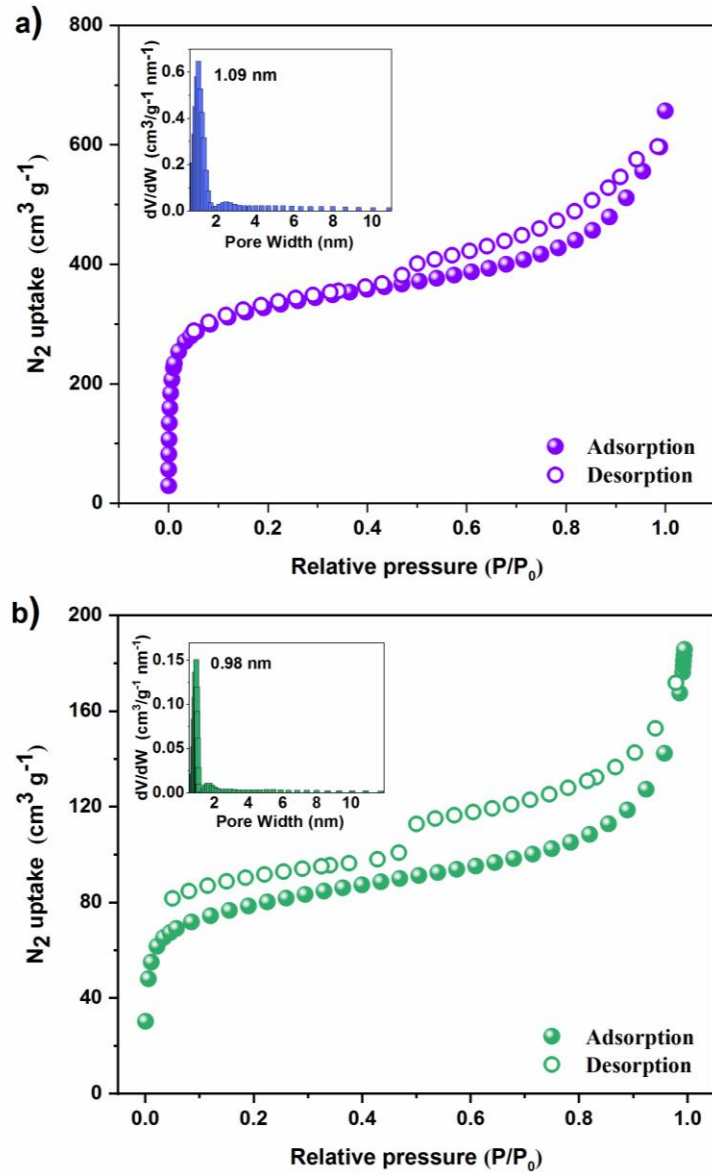


Figure 3. N₂ adsorption-desorption isotherms for JUC-570 (a) and JUC-571 (b) at 77 K. Inset: pore-size distribution calculated by fitting on the NLDFT model to the adsorption data.

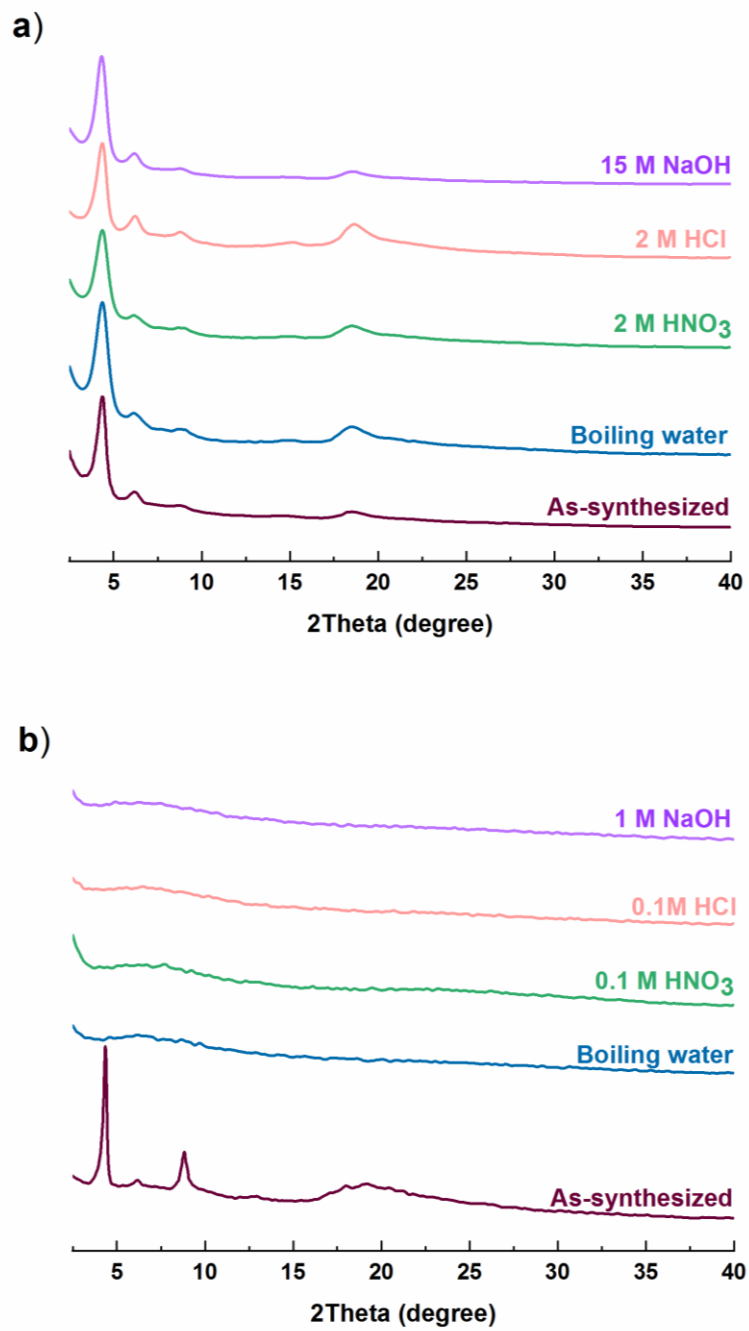


Figure 4. PXRD patterns of JUC-570 (a) and JUC-571 (b) after the treatment in strong acid, strong base and boiling water.

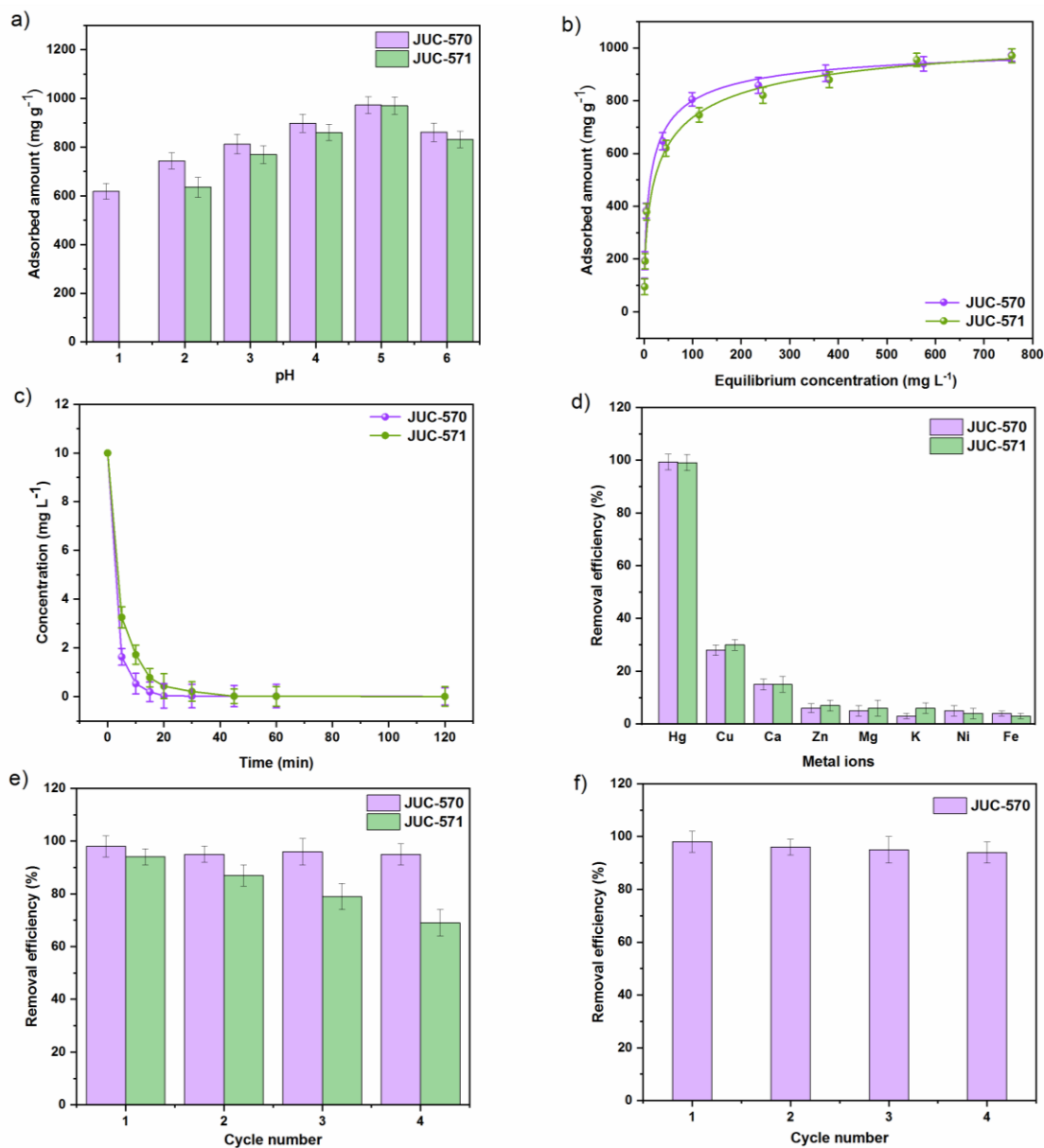


Figure 5. a) Effect of pH on the adsorption of Hg²⁺; b) Hg²⁺ adsorption isotherm of 3D thioether-based COFs at pH = 5 after 12 h; c) Hg²⁺ sorption kinetics of 3D thioether-based COFs under the initial Hg²⁺ concentration of 10.0 mg L⁻¹ and pH = 5; d) Selectivity test of 3D thioether-based COFs in the mixed ion solution; e) Recycling test of 3D thioether-based COFs at pH = 5; f) Recycling test of JUC-570 at pH = 1.

ToC:

Two 3D thioether-based COFs, JUC-570 and JUC-571, were for the first time prepared by the bottom-up approach, and used for mercury (Hg^{2+}) removal from aqueous solution. In particular, JUC-570 is periodically decorated with isopropyl groups around imine bonds, which displays higher chemical stability, BET specific surface area and Hg^{2+} removal performance.

Keywords: covalent organic framework, mercury removal, thioether functionalization

Yiying Zhang, Hui Li, Jianhong Chang, Xinyu Guan, Lingxue Tang, Qianrong Fang, Valentin Valtchev, Yushan Yan, and Shilun Qiu**

Three-Dimensional Thioether-Based Covalent Organic Frameworks for Selective and Efficient Mercury Removal