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Crystallographic and (spectro)electrochemical characterizations of cobalt(II) 10-phenyl-5,15-di-*p*-tolylporphyrin

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Abstract

The synthesis, cyclic and rotating disk electrode voltammograms, UV-visible absorption and X-ray diffraction analyses of cobalt(II) 10-phenyl-5,15-di-*p*-tolylporphyrin (**1-Co**) are described. **1-Co** was crystallized by slow diffusion of *n*-hexane into a concentrated CH₂Cl₂ solution. X-ray diffraction analyses reveals porphyrin aromatic cycle stacking in the crystal, C-H⋯π interactions of the CH₂Cl₂ solvent with the π-system of one tolyl group and Co(II)⋯π (porphyrin ring) interactions. The abstraction of 1.0 F/mol during the electrolysis at the first oxidation potential was followed by spectroelectrochemistry. It leads to the Co(II) → Co(III) transformation rather than the formation of the cation radical of the porphyrin ligand. These results were supported by comparison of the cyclic and rotating disc electrode voltammograms, of the UV-Visible absorption spectra, before and after electrolysis of **1-Co**.

Keywords

X-ray crystallographic structure; Cobalt, Porphyrin; Electrosynthesis; stacked aromatics dimers; C-H⋯π interactions; Co(II)⋯π interactions

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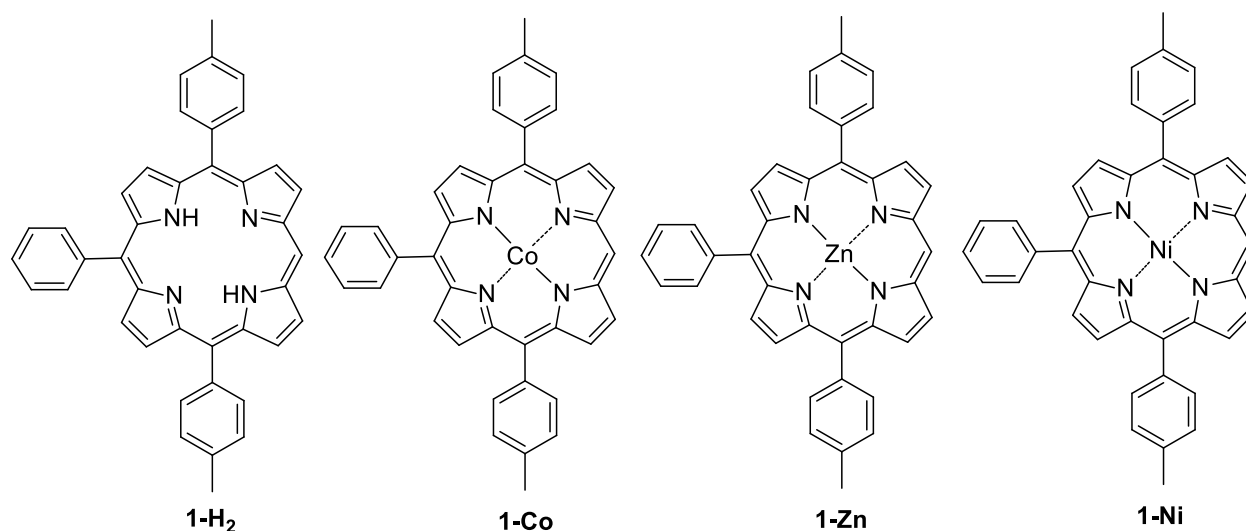
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1. Introduction

Porphyrins have been the subject of intense researches in numerous fields due to their essential role in natural processes such as photosynthesis and oxygen transport in blood. These macrocycles have found applications in material science, (electro)catalysis, energy converting devices and medicine [1-16]. Their properties can be easily adapted to specific applications by introducing new metal cations in the center of the porphyrin ring or by chemical functionalization of the porphyrin periphery [17, 18]. Cobalt porphyrins and corroles have attracted much attention in recent years, both because the cobalt metal exhibits several oxidation states and porphyrins can generally stabilize the central metal cations in their high oxidation states [19, 20]. Electrochemical four electrons reduction of O₂ to H₂O by cobalt porphyrins has been extensively studied so far not only because of its biological relevance, but also due to its technological significance such as in fuel cells [21, 22]. Furthermore, the studies of cobalamin and its derivatives/analogues have significantly contributed to increase the popularity of Co(II) complexes [23-26]. It was also found that the reactivity of axial ligation of cobalt porphyrin play a pivotal role in a number of bio- and 'bio-inspired' catalytic transformations. In 2018, Kadish and co-workers showed that equilibria involving the Co(III) derivative radical and the Co(II) porphyrin π -cation radical depends on the nature of a solvent and/or substituent [27-30]. In this context, it is important to know precisely the role of the various species involved in the mechanism and to control their reactivity. This requires a very fine description of each step, with the identification of the intermediates, the investigation of the reactivity of the electrogenerated species through cyclic and rotating disc electrode voltammetry analyses and preparative electrolyses. The work reported by Manassen elucidated the mechanism of the electrochemical oxidation of cobalt (II) *meso*-tetraphenylporphyrin (**CoTPP**) in CH₂Cl₂ by cyclic voltammetry [31]. In this report, it was clearly demonstrated that the first oxidation of **CoTPP** is centered on the metal (Co(II) \rightarrow Co(III)), while the two others correspond to the oxidation of the porphyrin ring.

As part of our on-going interest focused on the electrochemical reactivity of porphyrin complexes [15, 32-38], this work intends to study the electrochemical response and reactivity of cobalt(II) 10-phenyl-5,15-di-*p*-tolylporphyrin (**1-Co**, Scheme 1). Thus, in this manuscript, we report the synthesis, the electrochemical and X-ray crystallographic characterizations of **1-Co**.



Scheme 1. Free base porphyrin **1-H₂** [39] and its corresponding metal complexes **1-Co**, **1-Zn** [34] and **1-Ni** [35].

2. Experimental

2.1 X-ray equipment and refinement

Diffraction data were collected on a Nonius KappaCCD diffractometer equipped with a nitrogen jet stream low-temperature system (Oxford Cryosystems). The X-ray source was graphite-monochromated Mo- K_{α} radiation ($\lambda = 0.71073 \text{ \AA}$) from a sealed tube. Data were reduced by using DENZO software without applying absorption corrections; the missing absorption corrections were partially compensated by the data scaling procedure in the data reduction. The structure was solved by direct methods using the SIR92 [40] program and refined with full-matrix least-squares on F^2 using the SHELXL97 [41] program with the aid of the WIN-GX [42] program suite. All non-hydrogen atoms were refined with anisotropic thermal parameters. Hydrogen atoms attached to carbon atoms were included in calculated positions and refined as riding atoms. The molecular structure of **1-Co** has been determined by X-ray diffraction analyses. Crystal data, data collection and structure refinement details are given in Table 1. CCDC-2006293 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge via <https://www.ccdc.cam.ac.uk/structures/> (or from the Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033).

Table 1Crystal data and structure refinement for **1-Co**.

Empirical formula	2 C ₄₀ H ₂₈ CoN ₄ , CH ₂ Cl ₂
Formula weight	1332.11
Temperature	115(2) K
Wavelength	0.71073 Å
Crystal system	Triclinic
Space group	<i>P</i> -1
Unit cell dimensions	$a = 13.3891(3) \text{ \AA}$ $\alpha = 85.3960(10)^\circ$. $b = 15.1644(4) \text{ \AA}$ $\beta = 82.5310(10)^\circ$. $c = 15.5628(4) \text{ \AA}$ $\gamma = 88.6860(10)^\circ$.
Volume	3122.67(13) Å ³
<i>Z</i>	2
Density (calculated)	1.417 Mg/m ³
Absorption coefficient	0.673 mm ⁻¹
<i>F</i> (000)	1376
Crystal size	0.25 x 0.05 x 0.05 mm ³
Theta range for data collection	1.53 to 27.47°.
Index ranges	-17 ≤ <i>h</i> ≤ 16 -17 ≤ <i>k</i> ≤ 19 -20 ≤ <i>l</i> ≤ 20
Reflections collected	19132
Independent reflections	13942 [<i>R</i> (int) = 0.0541]
Completeness to theta = 27.47°	97.4 %
Max. and min. transmission	0.9671 and 0.8499
Refinement method	Full-matrix least-squares on <i>F</i> ²
Data / restraints / parameters	13942 / 0 / 842
Goodness-of-fit on <i>F</i> ²	1.196
Final <i>R</i> indices [<i>I</i> > 2σ(<i>I</i>)]	<i>RI</i> = 0.0803, <i>wR2</i> = 0.1649
<i>R</i> indices (all data)	<i>RI</i> = 0.1417, <i>wR2</i> = 0.1982
Largest diff. peak and hole	0.553 and -0.956 e.Å ⁻³

2.2 Reagents and instrumentation

Tetraethylammonium hexafluorophosphate (TEAPF₆, Fluka puriss., electrochemical grade, $\geq 99.0\%$) was used as received. CH₂Cl₂ (Carlo Erba 99.5%) and CH₃CN (SDS, Carlo Erba, HPLC gradient 99.9%) were distilled from P₂O₅ and CaH₂ respectively.

UV-visible absorption spectra were obtained with a Varian UV-vis spectrophotometer Cary 50 scan using quartz cells (Hellma). In spectroelectrochemical experiments, an UV-vis immersion probe (Hellma, $l = 2$ mm) was connected through two optical fibers to the spectrophotometer.

All electrochemical manipulations were performed using Schlenk techniques in an atmosphere of dry oxygen-free argon at room temperature ($T = 293 \text{ K} \pm 3 \text{ K}$). The supporting electrolyte was degassed under vacuum before use and then dissolved to a concentration of 0.1 mol L^{-1} . Voltammetric analyses were carried out in a standard three-electrode cell, with an Autolab PGSTAT 302 N potentiostat, connected to an interfaced computer that employed Electrochemistry Nova software. The reference electrode was a saturated KCl calomel electrode (SCE) separated from the analysed solution by a sintered glass disk filled with the background solution. The auxiliary electrode was a platinum wire separated from the analysed solution by a sintered glass disk filled with the background solution. For all voltammetric measurements, the working electrode was a platinum disk electrode ($\varnothing = 2$ mm). In these conditions, when operating in a mixture of CH₂Cl₂/CH₃CN 2/3 (0.1 M TEAPF₆) the formal potential for the Fc⁺/Fc couple was found to be +0.40 V vs. SCE.

Bulk electrolyses were performed in a three compartment cell with glass frits of medium porosity with an Amel 552 potentiostat coupled with an Amel 721 electronic integrator. A platinum wire spiral ($l = 53$ cm, $\varnothing = 1$ mm) was used as the working electrode, a platinum plate as the counter electrode and a saturated calomel electrode as the reference electrode. Electrolyses were followed by TLC and UV-visible absorption measurements.

2.3 Synthesis and crystallization

10-phenyl-5,15-*p*-ditolyl-porphyrin (**1-H₂**, see reference [31] for its preparation) (450 mg, 0.79 mmol) was dissolved under argon in 100 mL of CHCl₃/CH₃OH mixture (1/1 v/v) in the presence of 2.0 equivalents of cobalt acetate tetrahydrate and 2.5 equivalents of sodium acetate [43]. The resulting solution was then heated at reflux for 2h [30]. At the end of the reaction, the crude product

was washed three times with water (3 x 200 mL) then dried over MgSO₄. The organic phase was evaporated to dryness and the residue was purified on a silica plug (CH₂Cl₂ as eluent). The cobalt porphyrin was isolated in 70% yield (342.5 mg, 0.55 mmol). Suitable crystals for X-ray diffraction studies were obtained by slow diffusion of *n*-hexane into a CH₂Cl₂ solution of **1-Co**. MALDI-TOF MS (dithranol): [M]⁺ = 622.98; 623.16 calc. for C₄₀H₂₈N₄Co. UV-visible (CH₂Cl₂): λ max (nm) (ε, %) = 524 (100), 407 (6.3).

3. Results and discussion

3.1 Crystallographic structure of **1-Co**

The three-dimensional molecular view of **1-Co** is presented in Fig. 1. Two structurally different porphyrins (porphyrin **1** and porphyrin **2**) exist in the crystal. Both porphyrins exhibit a saddled-shape distortion and no axial ligand is coordinated on the cobalt(II) atom as observed by Martins and co-workers for **CoTPP**·(CH₂Cl₂)_{0.6} (CCDC 864256) [44]. In our case, one methylene chloride molecule per two porphyrins was found in the crystal lattice. Selected data are also given in Table 2 and are compared with two published structures of **CoTPP** without any coordinating ligands (CCDC 864256 [44] and CCDC 603708[45]). It must be stressed that when a coordinating molecule is present, it generally occupies the axial position of the cobalt porphyrin [28-29].

The RMS deviation for **1-Co** (0.212 and 0.203 Å, Table 2) is lower than those calculated for **CoTPP**·(CH₂Cl₂)_{0.6} (0.419 Å [44] and **CoTPP** (0.251 Å [45]). However, the distance between the metal cation and the mean plane of the porphyrin is quite similar for these cobalt complexes (0.002(2), 0.008, 0.000 Å for **1-Co**, **CoTPP**·(CH₂Cl₂)_{0.6} and **CoTPP**, respectively) but also in comparison with the corresponding **1-Zn** and **1-Ni** complexes (0.0100(15) and 0.007(3) Å, respectively, [34, 35]. For **1-Co**, the metal–N distance (1.953(4) < *d* < 1.966(4) Å) is quite similar to **CoTPP** (1.949(2) Å), **CoTPP**·(CH₂Cl₂)_{0.6} (1.924(5) < *d* < 1.967(6) Å) and **1-Ni** (1.936(5) < *d* < 1.942(5) Å) but shorter than observed for the zinc(II) analogue **1-Zn** (2.034(3) < *d* < 2.039(3) Å).

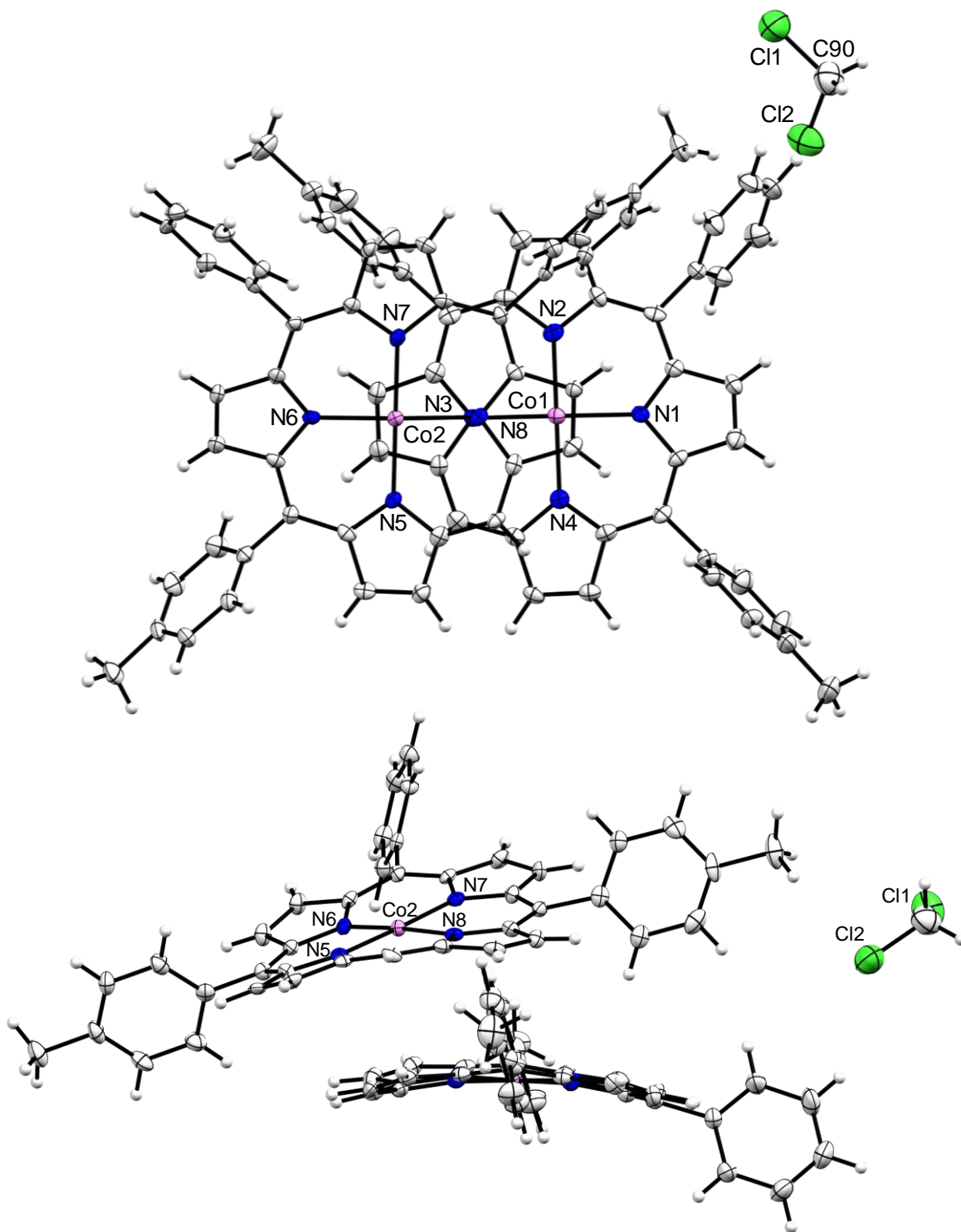


Fig. 1. Mercury views of **1-Co** X-ray crystallographic structure [Color code: C grey, N blue, Cl green, Co pink]. Thermal ellipsoids are scaled to the 50% probability level.

The torsion angle between the phenyl groups and the porphyrin plane is close to orthogonality in **CoTPP** ($80.21(10)^\circ$), presents some important variation in **CoTPP·(CH₂Cl₂)_{0.6}** (from

51.63(41)° to 83.74(14)°) and ranges from 62.87(13)° to 70.08(16)° for **1-Co**. For the tolyl groups of **1-Co**, the torsion angle with the porphyrin ring ranges from 61.85(14) to 74.98(11)° which are intermediate values as compared to the analogue zinc(II) complex **1-Zn** (from 84.59(9)° to 88.69(9)°, Table 2 [34]) and nickel(II) porphyrin **1-Ni** (from 66.69(17)° to 69.03(18)°, Table 2 [35]). The C_{meso}-C_{tolyl/phenyl} bond distances are intermediate between a single and a double bond (1.485(7) < d(C_{meso}-C_{tolyl/phenyl}) < 1.508(8) Å) and compare well with those measured for **CoTPP** (1.502(4) Å), **CoTPP·(CH₂Cl₂)_{0.6}** (1.495(7) < d(C_{meso}-C_{phenyl}) < 1.513(9) Å), **1-Zn** (1.501(4) < d < 1.504(5) Å) and **1-Ni** (1.479(8) < d < 1.500(8) Å).

The CH₂Cl₂ solvent used for the crystallization process interacts with both porphyrins *via* C-H···π interactions ($d(\text{C90-H90A}\cdots\text{centroid}(\text{C34-C35-C36-C37-C38-C39})) = 3.941(8)$ Å and $d(\text{C90-H90B}\cdots\text{centroid}(\text{C60-C61-C62-C63-N6})) = 3.769(8)$ Å) with the corresponding C-H···centroid angles equal to 150.3(5)° and 153.1(4)°, respectively (see Fig. 2). These values are in the range of those typically measured for this type of weak interactions [46].

Table 2Selected structural parameters for **1-Co**.

	1-Co	CoTPP·(CH₂Cl₂)_{0.6} [44]	CoTPP [45]	1-Zn [34]	1-Ni [35]
$d(\text{Metal-mean plane}_{\text{N}})^a$ (Å)	0.002(2) (porphyrin 1) / 0.019(2) (porphyrin 2)	0.008	0.000	0.0100(15)	0.007(3)
$d(\text{Metal-N})$ (Å)	1.953(4) < d < 1.966(4)	1.924(5) < d < 1.967(6)	1.949(2)	2.034(3) < d < 2.039(3)	1.936(5) < d < 1.942(5)
Ph torsion angles ^b (°)	70.08(16) (porphyrin 1) / 62.87(13) (porphyrin 2)	51.63(41), 83.74(14), 82.31(14), 82.38(15)	80.21(10)	76.23(12)	52.39(20)
Tol torsion angles ^b (°)	74.72(15), 68.87(15) (porphyrin 1) / 61.85(14), 74.98(11) (porphyrin 2)			84.59(9), 88.69(9)	66.69(17), 69.03(18)
$d(\text{C}_{\text{meso}}\text{-C}_{\text{Ph}})$ (Å)	1.498(8), 1.485(7) (porphyrin 1) / 1.508(7), 1.508(8) (porphyrin 2)	1.495(7), 1.499(7), 1.502(9), 1.513(9)	1.502(4)		1.501(4) < d < 1.504(5)
$d(\text{C}_{\text{meso}}\text{-C}_{\text{Tot}})$ (Å)	1.493(7), 1.488(8)				1.479(8) < d < 1.500(8)
RMS deviation ^c (Å)	0.212 / 0.203	0.419	0.251	0.0352	0.247
$d(\text{Co1}\cdots\text{Co2})$ (Å)	5.2296(10)	<i>e</i>	<i>e</i>	<i>e</i>	<i>e</i>
Mean planes _{porphyrin} Co1-Co2 angle ^d (°)	21.09(4)	<i>e</i>	<i>e</i>	<i>e</i>	<i>e</i>
density (g·cm ⁻³)	1.417	1.422	1.404	1.37	1.415

^a Mean plane_N: mean plane calculated with the 4 nitrogen atoms of the porphyrin core. ^b Angle between the mean plane_{porphyrin} and the mean plane of the phenyl and tolyl groups. ^c The RMS deviation corresponds to the root mean square of the distances of the 24 carbon and nitrogen porphyrin ring atoms from the mean plane formed by these atoms. ^d Angle between the two mean planes of the 24 carbon and nitrogen porphyrin ring atoms. ^e Not relevant.

Other C-H \cdots π interactions between tolyl groups and pyrrole rings from the porphyrins also contribute to stabilize the crystallographic structure (Fig. 3). Thus the distance $d(\text{C39-H39}\cdots\text{centroid}(\text{C55-C56-C57-C58-N7}))$ is 3.346(6) Å whereas distance $d(\text{C71-}$

H71...centroid(C11-C12-C13-C14-N2) is equal to 3.664(7). The corresponding C-H...centroid angles are equal to 152.1(4)° and 160.9(3)°, respectively.

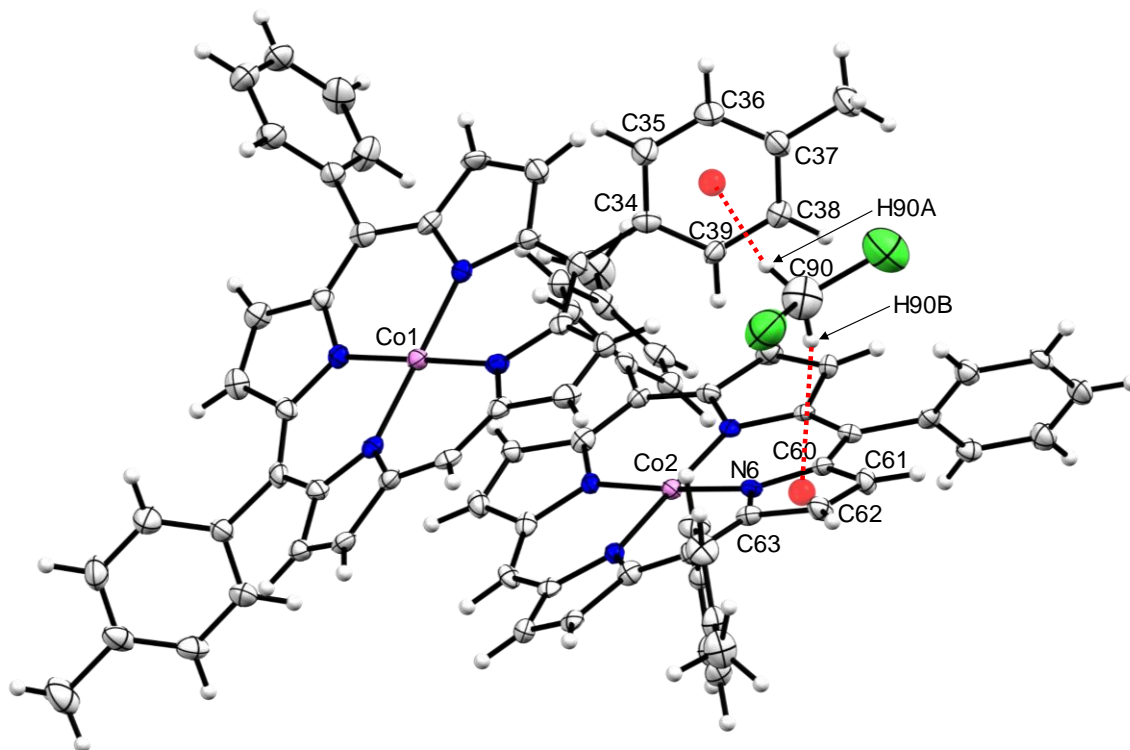


Fig. 2. Mercury view of **1-Co** presenting the C-H... π interactions of CH₂Cl₂ with the tolyl ring of porphyrin **1** and with the pyrrole ring of porphyrin **2** [Color code: C grey, N blue, Cl green, Co pink, centroids red]. Thermal ellipsoids are scaled to the 50% probability level.

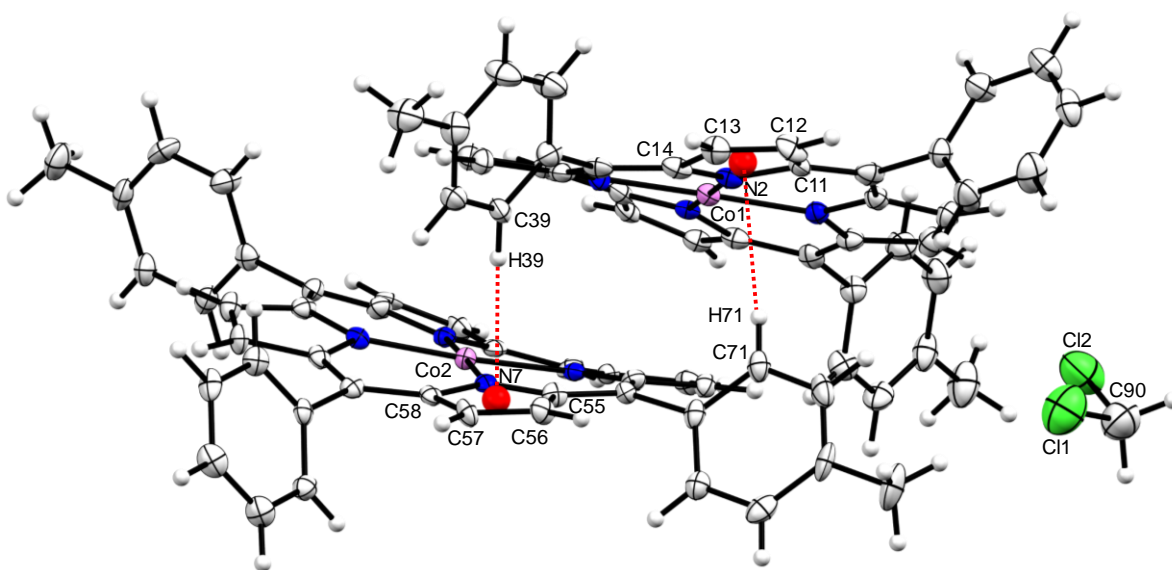


Fig. 3. Mercury view of **1-Co** presenting the C-H... π interactions of CH₂Cl₂ with the tolyl ring of porphyrin **1** and with the pyrrole ring of porphyrin **2** [Color code: C grey, N blue, Cl green, Co pink, centroids red]. Thermal ellipsoids are scaled to the 50% probability level.

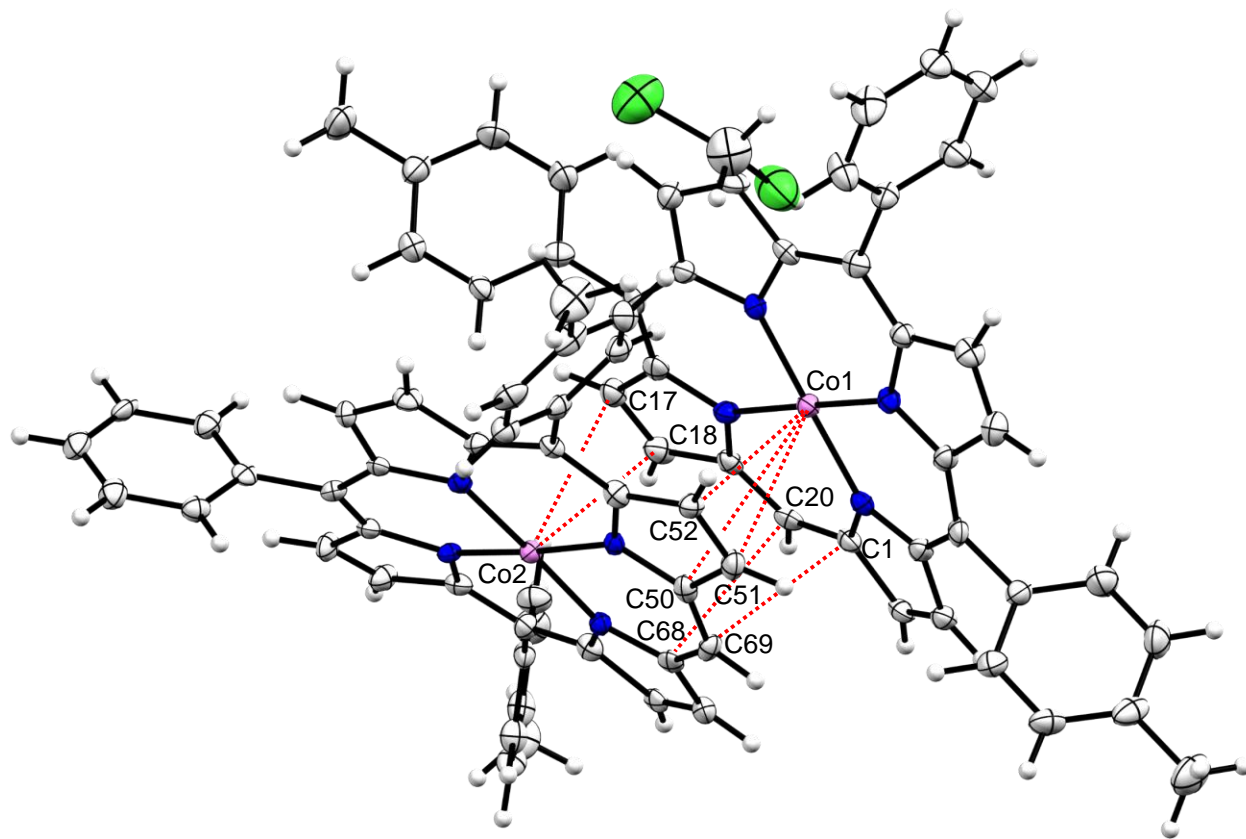


Fig. 4. Mercury view of **1-Co** presenting the Co(II)··· π and C(porphyrin 1)-C(porphyrin 2) interactions [Color code: C grey, N blue, Cl green, Co pink]. Thermal ellipsoids are scaled to the 50% probability level.

Remarkably, the Co(II) cation of the first porphyrin interacts with the π -system of the second porphyrin and *vice versa* (Fig. 4). This type of interactions has been previously reported for **CoTPP.(CH₂Cl)_{0.6}** [44]. For **1-Co** the distances between the cobalt(II) atom Co1 (from porphyrin 1) and the carbon atoms C50, C51 and C52 from porphyrin 2 are 3.661(5), 3.349(6) and 3.446(5) Å, respectively. Besides, the distances between the cobalt(II) atom Co2 (from porphyrin 2) and the carbon atoms C17 and C18 from porphyrin 1 are 3.489(6) and 3.664(6) Å, respectively. By comparison, for **CoTPP.(CH₂Cl)_{0.6}** the distances between the cobalt(II) atom Co1 and the carbon atoms C2 and C3 are significantly shorter (3.178(7) and 3.074(7) Å, respectively). The close porphyrin stacking is also strengthened by C(from porphyrin 1)···C(from porphyrin 2) short interactions with $d(\text{C68-C20}) = 3.274(8)$ Å and $d(\text{C69-C1}) = 3.278(7)$ Å.

Interestingly, **1-Co** molecules form three pairs of stacked aromatics dimers obtained by inversion symmetry centers along the a-axis with an angle between the two porphyrins included in the asymmetric unit equal to 21.09(4)° [47, 48]. Their slipped cofacial orientations resemble the

molecular arrangement found in the photosynthetic special pair [49]. Besides, the centroid-to-centroid (calculated from the 24 carbon and nitrogen porphyrin ring atoms) distances between porphyrins are equal to 5.230 (asymmetric unit), 4.319 (porphyrin **2** and its symmetrical) and 4.391 Å (porphyrin **1** and its symmetrical). Porphyrins are stacked in a non-parallel way along the a-axis, to form a pseudo-helical shape around the [1 0 0] direction (Fig. 5). The stacked porphyrin dimer formation might be favored by the absence of one *meso*-aryl substituent.

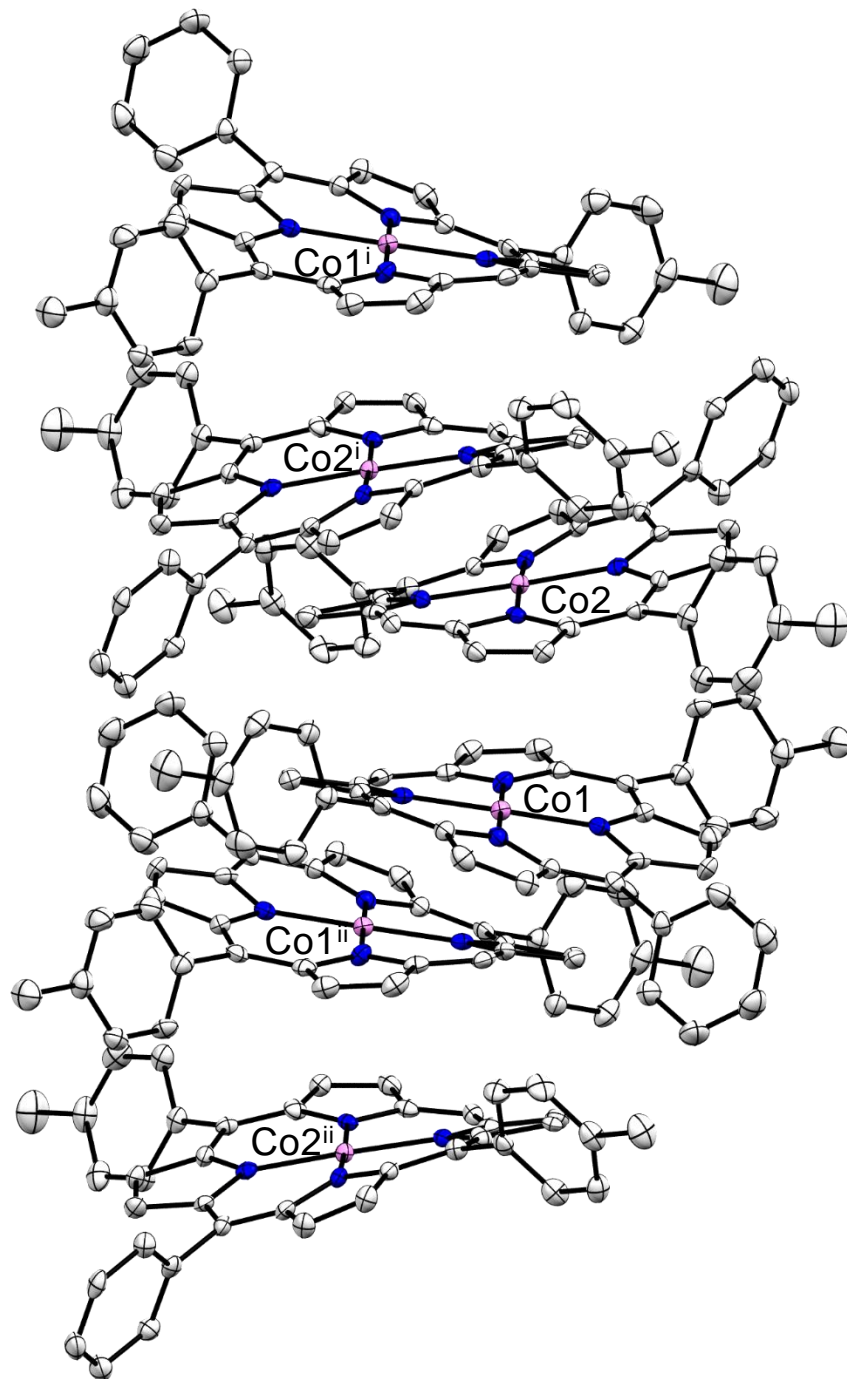


Fig. 5. View of the packing diagram of **1-Co** along the *c*-axis [Color code: C grey, N blue, Co pink]. Displacement ellipsoids are drawn at the 50% probability level. Hydrogen atoms were omitted for clarity. Symmetry codes : (i) $-x, -y, -z$; (ii) $1-x, -y, -z$.

3.2 Electrochemical characterization of **1-Co**

The cyclic voltammogram (CV) of **1-Co** in CH₂Cl₂/CH₃CN mixture (2/3 v/v) 0.1 M TEAPF₆ (Fig. 6) exhibits one monoelectronic irreversible metal centered oxidation (Co(II) → Co(III)) at $E_{pa}(O1) = 0.44$ V [31, 50]. The second redox process is reversible ($E_{1/2}(O2/R2) = 1.08$ V, $\Delta E_p = 70$ mV), but the third redox system is not ($E_{pa}(O3) = 1.49$ V). Previously, Manassen attributed the redox-systems that were involved in the electrochemical oxidation of cobalt (II) *meso*-tetraphenylporphyrin (**CoTPP**) [31]. Based on cyclic voltammetry investigations, these studies showed that the first oxidation peak was remarkably shifted to less positive potentials when coordinating solvents such as pyridine were added whereas other redox systems were less or not affected. Given the very close structural analogy of **1-Co** with **CoTPP**, it was initially assumed that the first oxidation of **1-Co** was centered on the metal (Co(II) → Co(III)) rather than on the porphyrin ligand. Thus, the two monoelectronic subsequent oxidations are ligand centered and correspond to the formation of the cation-radical (peak system O2/R2) and the dication (O3/R3), respectively.

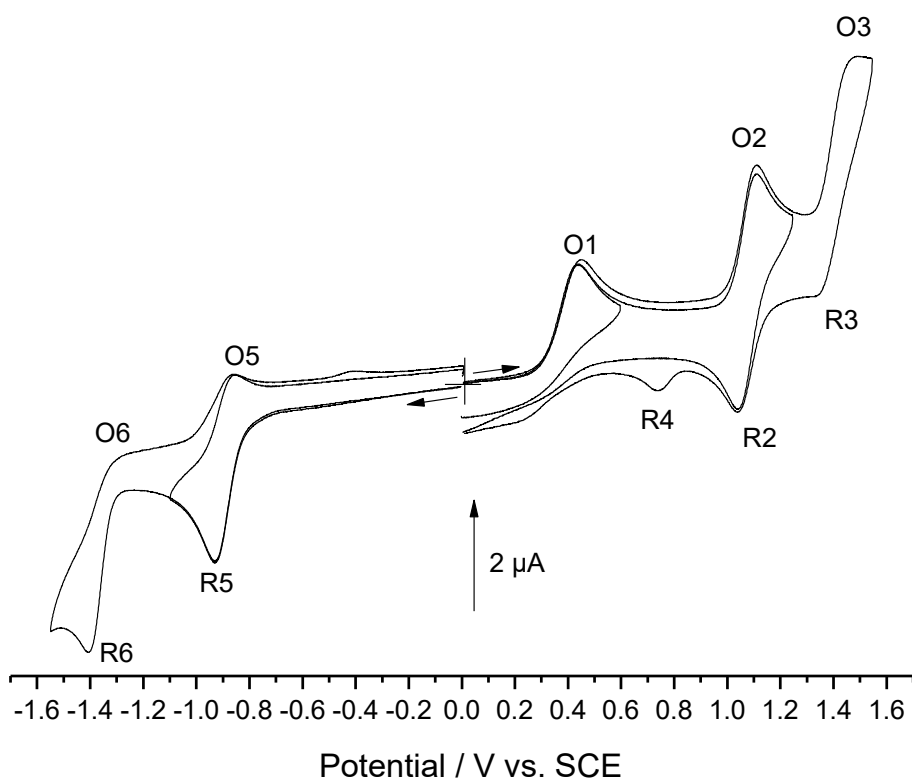


Fig. 6. Cyclic voltammograms of **1-Co** in CH₂Cl₂/CH₃CN (2/3 v/v) 0.1 M TEAPF₆ (WE: Pt, $\phi = 2$ mm, 0.1 V s⁻¹, [**1-Co**] = 5.0 × 10⁻⁴ M).

As summarised in Table 3, the potential of the oxidation peaks principally depends on the metal inside the porphyrin as well as its peripheral substitution pattern. In the case of cobalt(II) or nickel(II) porphyrins, the two monoelectronic oxidations peaks of the macrocycle (second and third oxidations for the Co(II) porphyrins and first and second oxidations for the Ni(II) porphyrin) are significantly higher than the zinc(II) complex. This evolution matches well with what was previously reported for CoTPP, NiTPP and ZnTPP [51].

Table 3

Potential values ($E_{1/2}$ or E_p for irreversible systems) for **1-Co**, **CoTPP**, **1-Zn** and **1-Ni** (in V vs SCE).

Compounds	Experimental conditions		Oxidation		
	<i>Solvent</i>	<i>Electrolyte</i>	1st	2nd	3rd
1-Co	CH ₂ Cl ₂ /CH ₃ CN (2/3 v/v)	0.1 M TEAPF ₆	0.44	1.08	1.49
CoTPP [29]	CH ₂ Cl ₂	0.1 M TBAP _a	0.75	0.98	1.16
1-Zn [34]	CH ₂ Cl ₂ /CH ₃ CN (4/1 v/v)	0.1 M TBAPF ₆	0.72	1.15	-
1-Ni [35]	CH ₂ Cl ₂ /CH ₃ CN (4/1 v/v)	0.1 M TEAPF ₆	1.03	1.25	-

When the positive potential scan is reversed back just after peak O3, a new reduction peak is seen at $E_{pc}(R4) = 0.74$ V. This additional R4 peak could be tentatively attributed to the reduction of an oxidized dimer that is formed during the forward scan [34, 35]. In the negative potential direction, the first monoelectronic pseudo-reversible peak ($E_{pc}(R5) = -0.93$ V, $\Delta E_p = 80$ mV) is attributed to the Co(II) \rightarrow Co(I) transformation by comparison with other data from the literature [52]. The CV exhibits a second irreversible reduction at $E_{pc}(R6) = -1.41$ V which plausibly corresponds to the first reduction of the porphyrin ring. To confirm that the first oxidation system was correctly assigned, UV-visible spectro-electrochemical experiments were performed under argon, at room temperature, in a CH₂Cl₂/CH₃CN mixture (2/3 v/v) containing 0.1 M of TEAPF₆ and **1-Co** at a concentration of $[1-Co] = 5.0 \times 10^{-4}$ M. The potential was set at a potential slightly more positive than peak O1 ($E_{app} = 0.60$ V). The current progressively decreased and finally reached the residual current value when 1.0 F per mol of **1-Co** was transferred. Then, the RDE voltammogram of the resulting red solution (Fig. 7) no longer shows the initial monoelectronic O1 oxidation system, which gives evidence for complete conversion of **1-Co**. The other monoelectronic oxidation processes (O2 and O3) are similar to the ones observed before electrolysis of **1-Co** (potential position and current magnitude).

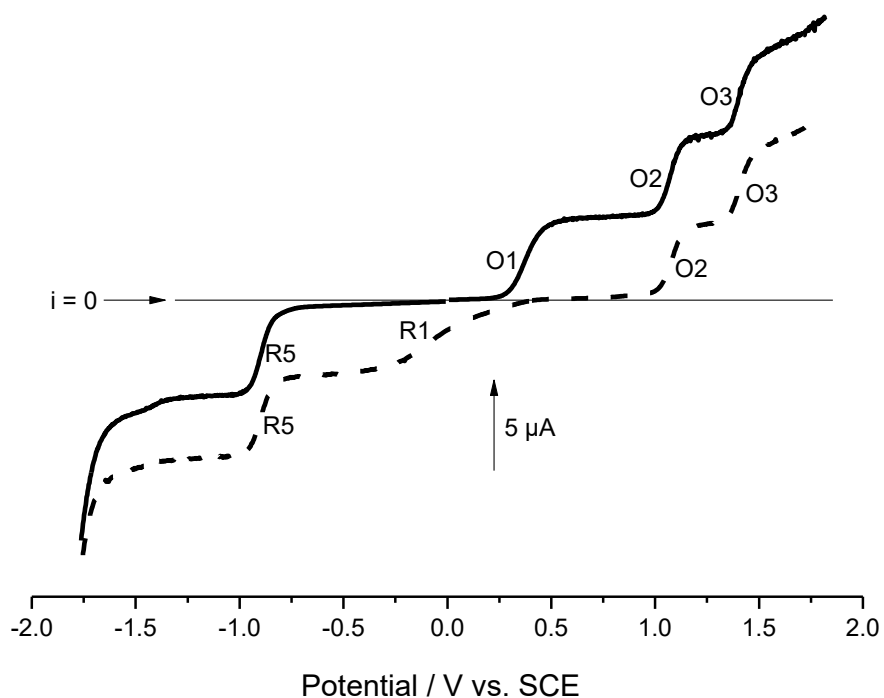


Fig. 7. RDE voltammograms of **1-Co** in $\text{CH}_2\text{Cl}_2/\text{CH}_3\text{CN}$ (3/2 v/v) 0.1 M TEAPF_6 , before (solid line) and after electrolysis (dashed line) (WE: Pt, $\varnothing = 2$ mm, 10 mV s^{-1} , $\omega = 500$ rpm, $[\mathbf{1-Co}] = 5.0 \times 10^{-4}$ M, - 1.0 F/mol of **1-Co**).

An additional cathodic system centered around 0.00 V is attributed to the $\text{Co(III)} \rightarrow \text{Co(II)}$ system which confirms the formation of **1-Co^{III}** after the electrolysis. The evolution of the UV-visible absorption spectra was simultaneously monitored during the electrolysis. A representative evolution is displayed in Fig. 8. The Soret band foot underwent a progressive bathochromic shift, whereas the initial Q band at $\lambda = 524$ nm progressively disappears at the expense of two new Q bands at $\lambda = 543$ and 576 nm. Another band is also observed at higher energy ($\lambda = 317$ nm). In another experiment, given the high molar extinction coefficient of the **1-Co** Soret band, the concentration was decreased to $[\mathbf{1-Co}] = 2 \times 10^{-5} \text{ mol L}^{-1}$ so that the maximum absorbance fell into the measurable spectrometer range. The initial Soret band of **1-Co** ($\lambda_{\text{max}} = 407$ nm) progressively disappears at the expense of a new band at 429 nm. The presence of isosbestic points located at 419 and 534 nm confirms the simple transformation of the initial porphyrin into another well-defined product, *i.e.* **1-Co^{III}**, with no intermediate observed between. During the first oxidation of **1-Co**, the UV-visible absorption spectra evolution was very similar to those previously reported for **CoTPP** by Wolberg and Manassen in benzonitrile (0.1 M $n\text{-Bu}_4\text{NClO}_4$) [53] and by Jones and Hinman in $\text{CH}_2\text{Br}_2/\text{CH}_3\text{CN}$ (0.1 M TBAPF_6) [54].

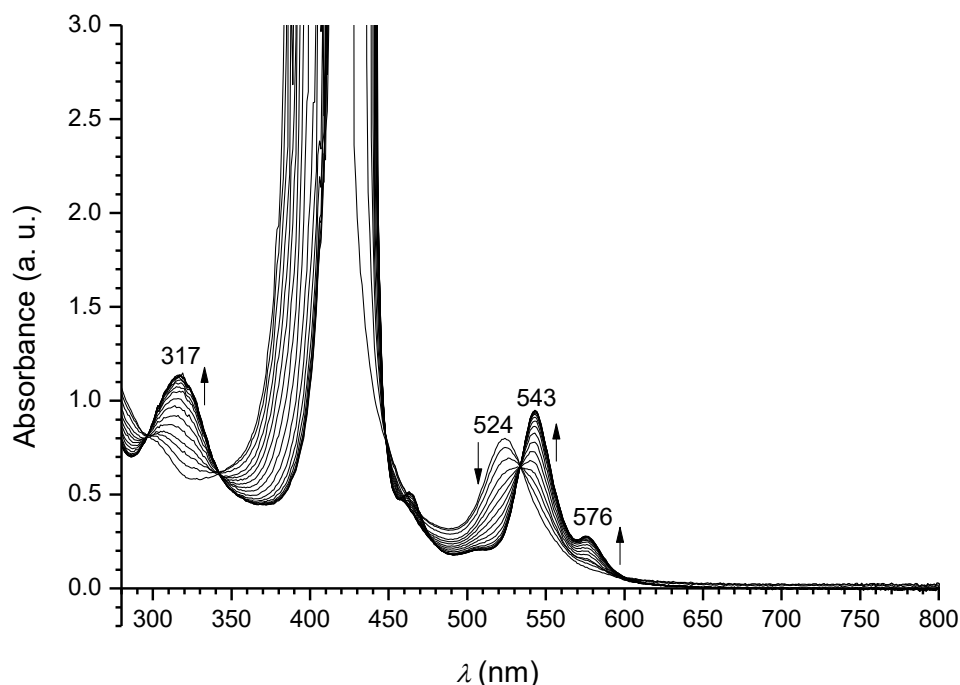


Fig. 8. Electrolysis of a 5.0×10^{-4} M solution of **1-Co** followed by UV-vis spectroscopy ($l = 2$ mm, 0.1 M TEAPF₆ in DCM/ACN 2/3 v/v, $E_{app} = 0.60$ V vs. SCE, -1 electron, WE: Pt wire, WE and counter-electrode are in separated compartments).

Based on EPR [55] and IR [54] analyses of the **CoTPP** first oxidation product, these authors demonstrated that this monoelectronic oxidation was reversible and centered on the metal and not on the porphyrin ring. Thus in our case the spectro-electrochemical experiment confirms again that this first oxidation concerns the reversible $\text{Co(II)} \rightarrow \text{Co(III)}$ transformation.

4. Conclusion

In summary, the first monoelectronic oxidation of **1-Co** was proved to be metal-centered ($\text{Co(II)} \rightarrow \text{Co(III)}$) based on electrochemical and spectroelectrochemical studies and also by comparison with the results reported in the literature. The original X-ray crystallographic structure of **1-Co** showed different kinds of interactions such as aromatic cycle stacking (stacked aromatics porphyrin dimers), $\text{C-H} \cdots \pi$ interactions between the dichloromethane molecule and the π -systems of tolyl and pyrrole rings, $\text{C-H} \cdots \pi$ interactions between tolyl groups and the π -systems of the porphyrins as well as $\text{Co(II)} \cdots \pi$ interactions between the cobalt(II) cations and the π -systems of the porphyrins. Also the stacked aromatics dimers arrange in a pseudo-helical manner. Our efforts are now focused on the electrochemical reactivity of **1-Co** at higher potentials.

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