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Hydroxy-channels – adaptive pathways for highly selective water cluster permeation

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ABSTRACT: Artificial water channels-AWCs are known to selectively transport water, with ion exclusion. Similar to natural porins, AWCs encapsulate water wires or clusters, offering continuous and iterative H-bonding that play a vital role in their stabilization. Herein, we report octyl-ureido-polyol AWCs capable of self-assembly into hydrophilic hydroxy-channels. Variants of ethanol, propanediol and trimethanol are used as headgroups to modulate water transport permeabilities, with complete rejection of ions. The hydroxy-channels achieve a single-channel permeability of 2.33×10^8 water molecules per second, which is within the same order of magnitude as the transport rates for Aquaporins. Depending on their concentration in the membrane, adaptive channels are observed in the membrane. Over increased concentrations, a significant shift occurs initiating unexpected higher water permeation. Molecular simulations probe that sponge-like or cylindrical aggregates can form to generate transient cluster water pathways through the bilayer. Altogether, the adaptive self-assembly is a key feature influencing channel efficiency. The adaptive channels described here may be considered as an important milestone contributing to the systematic discovery of artificial water channels for water desalination.

INTRODUCTION

One of the most ambitious goals in separation sciences today, is to accomplish the reconstruction of natural carriers or channels through synthetic design. Proteins have evolved for millennia to adopt their current functional structures. Although they are remarkable in functional efficacy, their structural complexity and environmental sensibility undermine their reproducibility and further applications for “out of the membrane cell” use.¹

A key step in achieving this goal is the replacement of natural transporters with simpler synthetic analogues. Synthetic biomimetic approaches, offer us the possibility to build artificial unimolecular or supramolecular constructs that can mimic biological functionality, while featuring procedural and structural simplicity.² The functions of carriers, ion-channels and porins have been successfully replicated using small building blocks. There are several advantages for using artificial mimics; most notable are the possible lower costs, tunability, improved stability and compatibility with the membrane host matrices.

After a decade of their discovery,³⁻⁷ Artificial water channels – AWCs come very close to the performance of the natural counterparts, the Aquaporins. Carbon Nanotube Porins-CNTPs (Figure 1a,b),^{8,9} Pillar[5]arene-PAP5 (Figure 1c)^{10,11} Pillar[4]arene-PAP4 clusters (Figure 1d),¹² Aquafoldamers (Figure 1e),¹³ Porous organic cages-POC (Figure 1f),¹⁴ Pillar[5]arene Hydrazones-PAH5⁷ and Pillar[5]arene-AQPs¹⁵ are able to close- or even out-perform AQPs, when it comes to water permeability and/or selectivity.¹ In all artificial setups, transport of water occurs either through a single water file, as seen in the

case of AQPs, or through the water cluster structures inside or outside the cavity of the channels.¹⁶

Increasing water-salt permselectivity or improving the salt rejection is the actual challenge to boost the efficiency of AWCs.¹⁶ Most artificial channels lack the complex mechanism involved in selective water transport, including wetting-dewetting events, intermolecular water-water or channel-water H-bonding, all these translated to water translocation through the channels. Only a few synthetic channels (I-quartet, PAP4, POC, Aquafoldamers, Pillar[5]arene -AQPs) are capable of selective rejection of ions. In these examples, the water transport takes place *via* water wires or clusters, through a protective ion-exclusion mechanism by the channel confined water superstructures.^{1,16} This mechanism imposes an optimal feeding of water-rich wires/clusters, that selectively form H-bonds to amphoteric water and reject ions or other molecular solutes.^{17,18}

Structural issues with channels, dimensionally adapting the formation of wires or clusters, depending on their interior pore diameter, can enhance/reduce the water permeability that can significantly impede/exceed the channel selectivity. Weak interactions of water molecules with the hydrophobic walls of channels contribute to its nearly frictionless transport, while selectivity is sterically controlled. Structural design represents a formidable challenge in synthetic chemistry of AWCs. Within this context our pioneering work on I-quartet^{4,17} (Figure 1i) or I-quartet sponges¹⁸ (Figure 1j) highlights the tremendous importance of controlling the water superstructures *via* channel-water H-bonding events between communicating water molecules in the channel cavity, exerting the prohibition of ion insertion/translocation events. Obtaining absolute water

selectivity resides in designing a channel, at least in its selectivity filter/region, with an inner binding environment that is favorable for the selective recognition of water *via* synergetic donor-acceptor H-bonding and ion exclusion.

Basically, one molecule of water is able to form four hydrogen bonds, simultaneously acting as donor and acceptor through two bonds for each. We previously observed near perfect single water-wire formation within I-quartet channels,^{4,17} where the inner walls of the channels interact with a water molecule through one donor and one acceptor hydrogen bond, leaving the two other H-bonds to be used to interconnect the water wires throughout the channel. Using the same principles, highly specific unimolecular water channels have been achieved using Aquafoldamers¹³ or Pillar[5]arene-AQPs¹⁵ in which water wires are stabilized *via* channel donor-acceptor H-bonding, while having a lower performance than the AQPs and CNTP.^{2,3}

Increasing the water-water H-bonding results in the formation of clusters, that become highly selective in interactions with other species that are mostly rejected as they perturb the stability of water superstructures. Within this context, the concept of engineering cluster/sponge-type AWCs is clearly evolving, as it considerably expands the range of possibilities for the design of macroscopic biomimetic membrane materials embedding AWCs for reverse osmosis desalination.¹⁹

Combined with our previous experience in constructing self-assembled AWCs,^{17,18} herein we report a new class of AWCs based on octyl-ureido-polyols **H1-H6** (Scheme 1) bearing hydrophilic hydroxyl groups, that form adaptive channels in lipid bilayers. Accommodating single water-wires or water-clusters, the channels are able to selectively transport water across a lipid membrane. We argue that such selective transport occurs through a selective water channel/pore assisted diffusion mechanism, controlled by donor-acceptor hydroxyl bonding groups substantially influencing their permselectivity.

This scheme underlines the functional versatility of the artificial OH-channels, reminiscent to ceramide phospholipids that are now well-known for more than 40 years to be involved in cell regulation (apoptosis of cancer cells, differentiation...),^{20,21} Beside their cell regulation properties, ceramides are also known, when mixed with phospholipids, to generate two main effects: they increase the molecular order of phospholipids and they give rise to lateral phase and domain formation. Indeed, increase of the membrane permeability has been studied by Colombini and al. that have demonstrated the formation of pores.²² To the best of our knowledge, ceramides are the only class of non-proteogenic natural compounds able to form supramolecular assemblies that span the membrane bilayer, to generate pores for facilitated ion/water translocation.

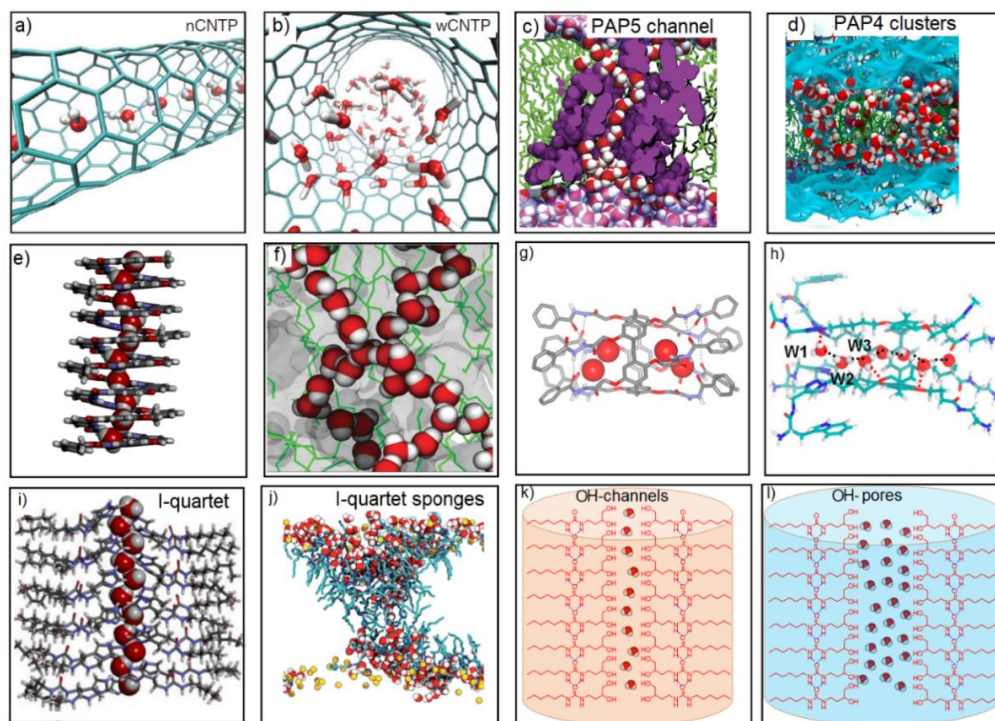


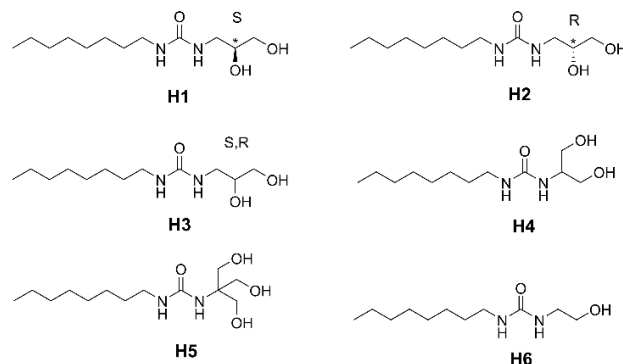
Figure 1. Artificial water channels. a) Narrow *n*CNTPs and b) wider *w*CNTPs Carbon Nanotubes Porins encapsulating single water-wires and water clusters;⁸ c) Pillar[5]arene-PAP5 channels for water-wires transport;¹¹ d) water permeation through PAP4 clusters in bilayer membrane;¹² e) single water-wire permeation through highly selective Aquafoldamers;¹³ f) water chains/clusters within POC nano-aggregates embedded in bilayer membrane;¹⁴ g) encapsulated water dimers within Pillar[5]arene Hydrzones-PAH5;⁷ h) Molecular structure of Pillar[5]arene-AQPs¹⁵ and its interaction with the single water-wire file; i) I quartet water channels and their interaction with single water wires;¹⁷ j) water clusters within I-quartet sponges embedded in bilayer membranes;¹⁸ adaptive k) narrow OH-channels and l) wider OH-pores described in this work, presenting hydroxylic decorated pathways for water-wires and water cluster translocation, through directional O-channels or OH-pores in bilayer membranes.

RESULTS AND DISCUSSION

Design Strategy. Herein, we focused on using simple octyl-ureido-polyols building blocks **H1-H6** (Scheme 1), that would be capable of combining behaviors of the I-quartet and natural ceramide channels. Therefore, we synthesized a series of compounds containing water/ion hydroxylic interacting heads and bearing an octyl chain in order to introduce optimal lipophilicity.¹⁷ The connection between these two elements is one urea group, known to form directional ribbon-type superstructures.¹⁸ Hydroxyl, as the most simple organic group for forming both donor and acceptor hydrogen bonds, has a natural affinity with water molecules. Ethanol, propanediol and trimethanol were employed as the hydrophilic binding motif for obtaining selective AWCs. Similar to the imidazole groups in I-quartet^{17,18} the “amphoteric” hydroxyl groups can selectively bind the water molecules through the donor-acceptor H-bonds.

Six compounds were synthesized for the studies presented here (Scheme 1). They were obtained through a one-step reaction between octyl-isocyanate and different amino-alkyl-polyols in a mixture of solvents. The ¹H-, ¹³C-NMR and ESI-MS spectra agree with the proposed formulas (Figure S1-S18). The potential to form OH-channel superstructures was assessed initially by crystallization of compounds **H1-H6**. X-ray single-

crystal structures give a good indication of the self-assembly preferences of the compounds in solution and their possible self-organization in bilayer membranes. However, these assumptions may be treated with caution, as the solid-state single-crystal structures are representing only snapshots of possible arrangements of functional channels or pore forming aggregates in the bilayer membrane.



Scheme 1. Compounds **H1-H6**, as components for self-assembled hydroxy OH-Channels.

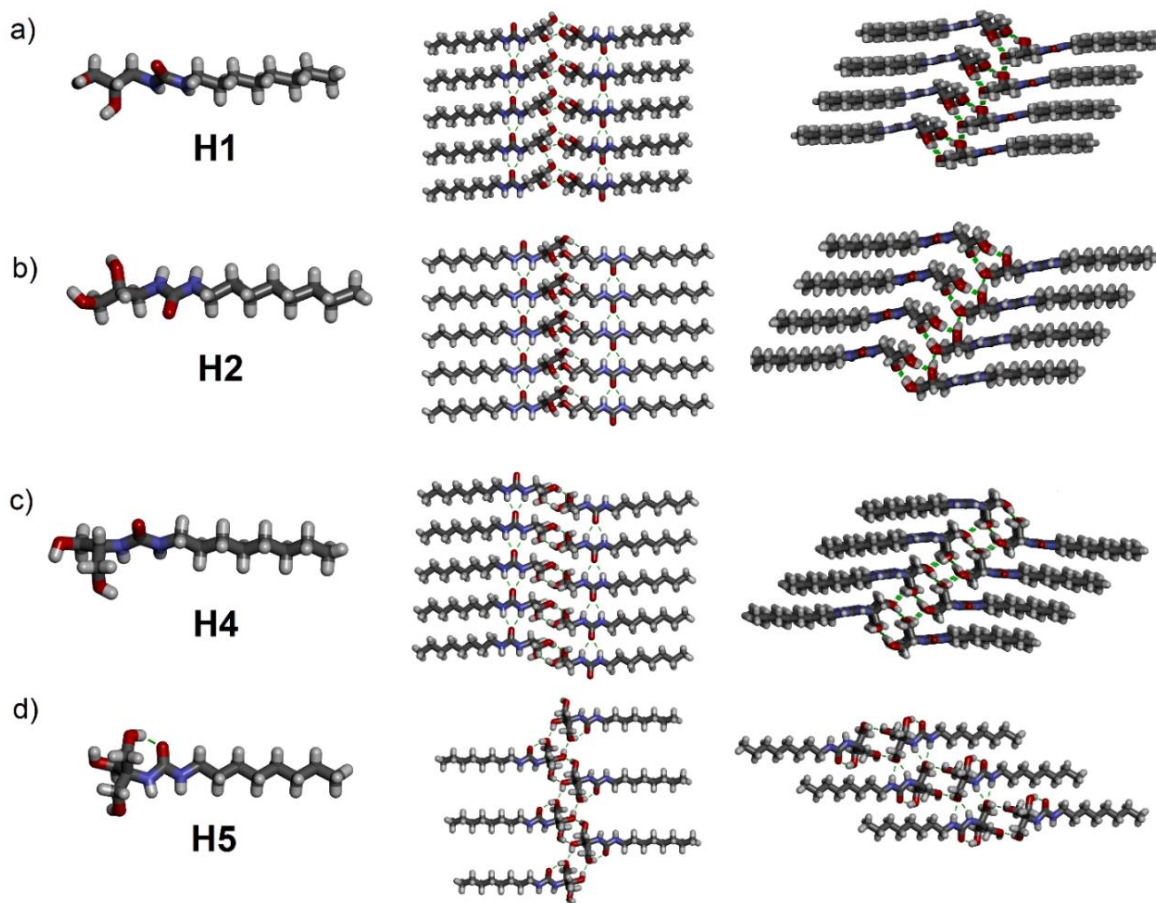


Figure 2. Crystal structures in stick representation of : a) **H1**, b) **H2**, c) **H4** and d) **H5** (N, blue; C, gray; O, red; H, white) showing (left) the monomeric molecules together with the lateral packing of H-bonded hydroxyl tetrads (middle) and their perpendicular packing via urea-urea and urea-hydroxyl H-bonding ribbons (right).

Single-crystals of good quality for X-ray structure determination were obtained for 1,2 diols **H1** and **H2**, 1,3 diol **H4** and 1,3,5 triol **H5** (Table S1). Their solid-state structures reveal the formation of networks of (OH---O)₄ H-bonding quartets, with the participation of two (**H1**, **H2** and **H5**) and four (**H4**) different molecules each involving two or one hydroxyl arms, respectively. They form almost perfect rectangular assemblies with the sides of ~2.7 Å. This OH tetrameric motif, then repeats both perpendicularly, as quartet component molecules are stacking through the urea moiety in crystal structures of **H1**, **H2**, and **H4**, as well as laterally, the second arm of the compounds participating in the formation of another adjacent OH-quartet. The distance between laterally repeating units is ~3.5 Å (Figure 2a-c). The double H-bonding between the third HO-group and adjacent HNCONH urea groups disrupts the formation of urea ribbons, but preserves a strong perpendicular H-bonding between interconnected OH-quartets in the crystal structure of **H5** (Figure 2d). Figures S19-S22 in Supporting information present structural details and CCDC 2015548-2015551 contains the supplementary crystallographic data for this paper.

Bilayer membrane transport experiments. In order to determine the transport permeability and selectivity, the OH-channels **H1-H6** were tested for water and ion transport experiments. Water transport was quantified through a scattering modulation method of a vesicle suspension using a stopped-flow setup,¹⁷ registering the shrinkage rate of vesicles under osmotic pressure to determine the induced osmotic water permeability of the OH-channels, in the presence of D(+)-sucrose or NaCl, as competitive osmolytes (Figure 2). To determine if the compounds are also inducing ion transport, we conducted pH sensitive HPTS fluorometric experiments for different cations (Li⁺, Na⁺, K⁺, Rb⁺ and Cs⁺) and anions (Br⁻, I⁻, HCO₃⁻, CO₃²⁻, NO₃⁻ and SO₄²⁻) and following the flow of species through the OH-channels. By using control FCCP and Valinomycin as an assisting H⁺ and K⁺ carrier, respectively, we were able to check the specificity of the OH-channels to transport ions and to determine the intrinsic selectivity of the channels, based on deconvoluted transport rates for cations and anions separately, rather focusing on their proton transport behaviors that could kinetically influence the translocation process. The transmembrane experiments offered all the prerequisites to assume that the compounds **H1-H6** assemble into functional and highly selective water-permeable OH-channels in liposomes.

Water-Transport Experiments. The OH-channels were reconstituted into phosphatidylcholine (PC) / phosphatidylserine (PS) / Cholesterol (Chl), PC/PS/Chl mixture with a molar ratio of 4/1/5 lipid vesicles (100 nm in diameter). The channel-forming compounds were added by simple addition of the dimethyl sulfoxide DMSO solutions of **H1-H6** to pre-formed vesicle suspensions using different compound to lipid molar ratios-mCLR (see Supporting information for details). Then, the vesicles were exposed to outwardly directed osmotic pressure gradients. Under hypertonic conditions driven by outwardly 10 mM PBS buffer solution (pH= 7.4) of 200 mM NaCl or 200 mM D(+)-Sucrose osmolytes (Figure 3, S24, S25), the shrinkage of the liposomes increased the light-scattering signal (Figure 2a). The water transport rates are strongly dependent on the concentration of the hydroxy-channels, that increase the overall water permeability relative to the background lipid permeability in the bilayer membrane. The net permeabilities show a continual increase when mCLR increase. A large range of

permeabilities from 9.36 μm/s to 19.63 μm/s can be obtained at mCLR < 1.48 for **H1-H6**. Interestingly for **H2**, **H3**, and **H4** we observe an abrupt 10-fold increase in water permeability values (160 to 560 μm/s) for mCLR > 1.48 (Figure S24, S25), reminiscent with the formation of larger channels/pores within the membrane. This unique adaptive behavior is probably related to an important concentration dependent morphology change of more permeable channels presenting different self-assembly behaviors within the membrane. To confirm this hypothesis the water transport measurements (Figure 3) were performed under different hypertonic solutions containing competitive osmolytes D(+)-sucrose (Figure S24) or NaCl (Figure S25) covering low and high channel concentrations domains.

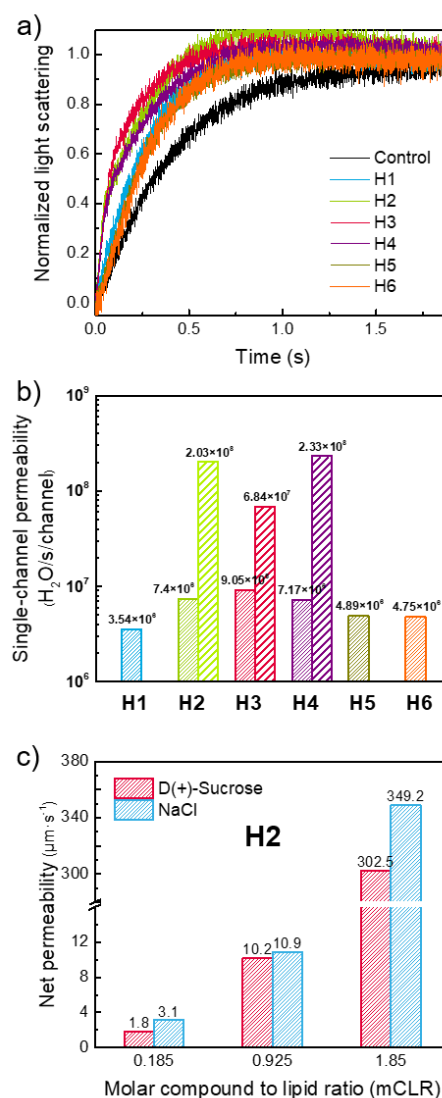


Figure 3. Measurements of water permeability through OH-channels. a) Stopped-flow light scattering traces of liposomes containing channels at different mCLR = 1.85, 1.67, 1.85, 1.67, 1.85 and 1.67, respectively; b) Single channel permeability calculated from the light scattering traces at different concentration in the membrane; c) The comparison of net permeability values for channels **H2**, determined by stopped-flow light scattering experiments with D(+)-Sucrose (pink) and NaCl (blue) as osmolytes.

It can be noticed that with D(+)-sucrose or NaCl the single channel permeabilities remain on the same level of magnitude at lower concentration for all channels and show a tendency of P_s increasing for **H2** and **H4**, while decreasing for **H3**, **H5** and **H6** with NaCl osmolarity compared with D(+)-sucrose osmolarity for the high concentrations. The channel-selection of D(+)-sucrose *via* size exclusion, may water selectively flux through the for **H3**, **H5** and **H6** channels, while the presence of NaCl osmolyte as the competitive salt transport agent might hinder the hydrophilic hydroxyl groups to remain accessible to water. The situation is completely opposite for **H2** and **H4** showing higher water to salt selectivity at high concentration

We then performed carboxyfluorescein (CF) leakage experiments²³ to monitor channel formation at low and high concentration of **H1-H6** (Figure S26). The magnitude of CF release is clearly different and the activity is increased in the following order: **H4**>**H3**>**H6**>**H1**> **H2**=**H5**. According to these CF leakage experiments, it is possible to say that, at a high concentration for **H4**, **H3** and unexpectedly for **H6**, there is the formation of large pores that allow efflux of CF, while no activity was observed for **H1**, **H2**, **H5**. Colombini et al.²² observed the same behavior with ceramide, showing experimentally that the average channel size increased with the increase of ceramide content in the membrane.

Table 1. Single-channel permeability of **H1-H6** channels

OH-Channel	Single-channel permeability water molecules/s
H1	3.54×10^6
H2	2.03×10^8
H3	6.84×10^7
H4	2.33×10^8
H5	4.89×10^6
H6	4.75×10^6

We further estimated the single-channel permeability, P_s of **H1-H6**, based on the OH-channel insertion, lipid concentration and channel configuration in lipids (Table 1)¹⁸ In order to determine the single-channel permeabilities, we used two rational assumptions: i) spontaneous total insertion of the compounds into the lipid bilayer. The total partitioning coefficients of **H1-H6** in the bilayer membrane are confirmed as at maximal concentration of 80 mM solutions are perfectly clear in the presence of lipid vesicles, while precipitate turbidity is formed when the compounds are added to aqueous solutions, reminiscent with almost total partition of **H1-H6** in the bilayer membrane (please note that lower partition coefficients would lead to higher single channel permeabilities than those considered for total partition and ii) number of molecules needed to form a channel that spans the bilayer equal to the number of molecules found in the single crystal structure that correspond to 8 stacks of **H1-H6** OH-quartets and 64 molecules per assembly within a transmembrane length of 5 nm. Accordingly, the single channel permeability values varied obviously, covering the range of 3.5×10^6 - 2.4×10^8 water/s/channel at CLR 0.026, which are only one orders of magnitude lower than that of AQP ($\sim 10^8$ - 10^9 water /s/channel^{1,2}) (Figure 3, Table 1).

Water transport activities increase substantially at higher concentrations of **H2-H4**, implying the formation of highly permeable channels. Indeed, the hydrophilic diol heads seemed to be more optimized for permeation than monol and triol heads, leading to higher water permeation when mixed with lipids. The tendency for increase of net and single channel permeability at high concentration can be explained by progressive higher cluster-type aggregation of OH-channels, promoting the increased water transport via sponge like structures as we previously observed for I-quartet channels.¹⁸ Differently, a low concentration of OH-channels prevents the aggregation and allows only reduced water-wires translocation.

Cation and Anion Transport Experiments. Cation ($M^+ = Li^+ Na^+, K^+, Rb^+$ and Cs^+) transport activities across the bilayer membranes incorporating **H1-H6** OH-channels, reconstituted into phosphatidylcholine (PC) lipid vesicles (100 nm), at the same mCLRs as used for the water transport, were assessed using standard HPTS fluorescence assays (Figure S27, S28).^{24,25} Indeed, when tested with MCl on external buffer, the **H1-H6** channels do not present a dose-response-type activity, showing constant near to zero activity behaviors, independently of low (10 mM, Figure S27) or high (80 mM, Figure S28) concentration of active channels in the membrane. It is noteworthy that cations are not transported and channel structuration is no effective to induce M^+/H^+ antiport conductance states along the assembled channels even for **H2-H4** presenting an important switch in water permeability at high concentration of 80 mM.

We were then interested in supplementary evidence on the anion translocation mechanisms through the H-bonded hydroxyl networks, already used as channel recognition motifs for anion/proton co-transport.^{26,27} In order to check the anion transport activity resulting from anion/chloride antiport or anion/ H^+ co-transport, we conducted the HPTS assays. Different sodium salts (Br^- , I^- , HCO_3^- , CO_3^{2-} , NO_3^- , SO_4^{2-}) were used to determine the anion transport activity of **H1-H6** OH-channels. On this basis, the resulting transport profiles showed rather weak near to zero transport rates for the different NaX salts in the presence of 10 mM (Figure S29) and 80 mM (Figure S30) of **H1-H6** OH-channels, indicating that the transport of the hydrated anions is not governed by the solvation free energy of the anions. (Figure S30). Compounds **H1** and **H6**, showing low water transport activities, present very low (<5%) transport activity; most probably related to a Cl^-/X^- ($X^- = Br^-, I^-, NO_3^-, SO_4^{2-}$) antiport. In order to verify if the H^+ transport is the rate-limiting step for M^+/H^+ antiport activity, we conducted HPTS assays with compounds **H1-H6** in the presence of carbonyl cyanide-4-(trifluoromethoxy)-phenylhydrazone, FCCP as a H^+ carrier (Figure S28).²⁵ These experiments have not shown a difference with the previous results. All tested OH-channels are totally inactive for cation transport in the presence or the absence of FCCP proton carrier. Moreover, we were interested in further testing the compounds specificity by testing them for proton transport. We thus used a classical HPTS assay and control valinomycin (K^+ cations) and FCCP (proton) transporters, known to generate a strong K^+/H^+ electrogenic antiport (Figure 4c, d). Potential proton transport acceleration is unlikely given in the case of coupling of valinomycin with **H1-H6**, suggesting very low proton transport rates through OH channels in HPTS assays.

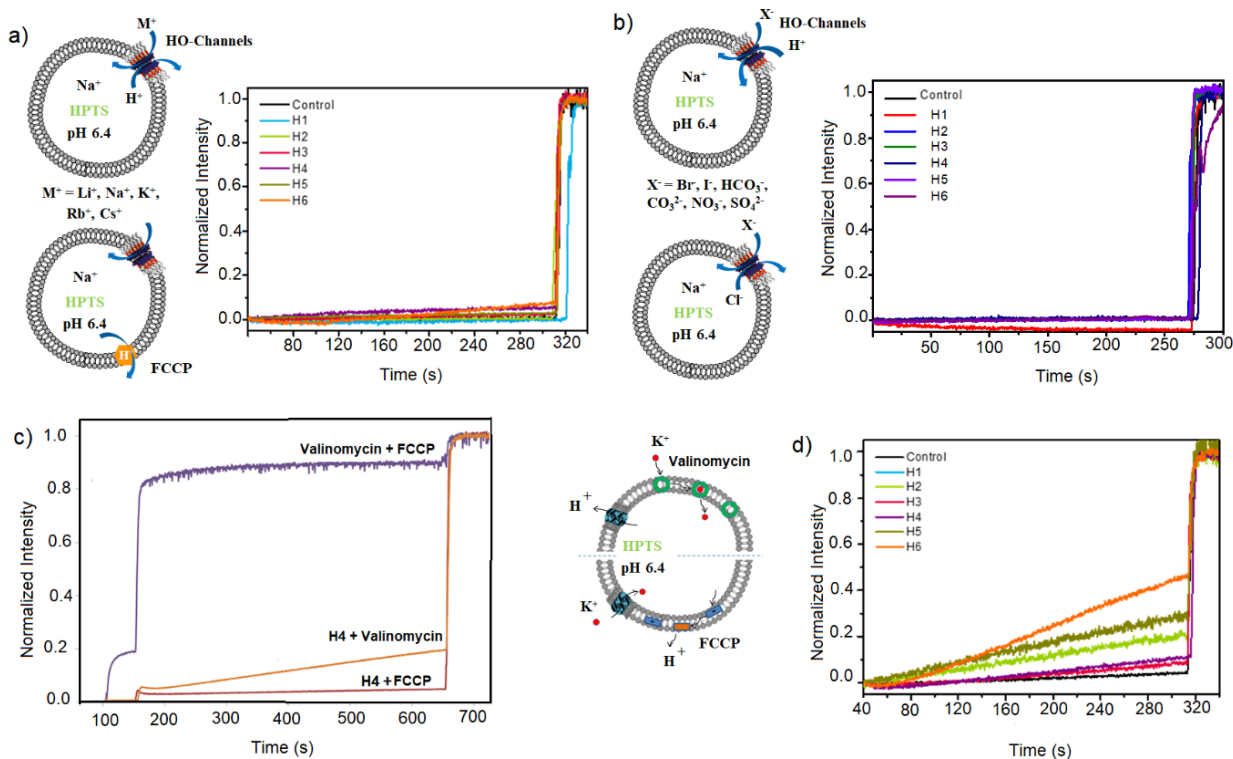


Figure 4. a) Comparison of Na^+ transport activity expressed as normalized fluorescence intensity of 10 mM channels in the extravesicular media containing 100 mM NaCl (in the absence of FCCP) together with schematic representation of antiport M^+/H^+ mechanisms in the absence or with the presence of FCCP as H^+ carrier. b) Comparison of Br^- transport activity expressed as normalized fluorescence intensity of 10 mM H1-H6 channels in the extravesicular media containing 100 mM NaBr, together with schematic representation of anion/ H^+ symport and Cl^-/X^- antiport mechanisms. c) Transport experiments coupling **H4** channels with valinomycin and FCCP as K^+ and H^+ carriers. d) Coupled transport between valinomycin K^+ carrier (50 μM) and **H1-H6** channels (80 mM) - 100 mM K_2SO_4 in the extravesicular media.

Molecular dynamics simulations of H1 aggregates in a fully hydrated lipid bilayer environment. Molecular dynamics simulations carried out in a membrane environment, allow us to observe a range of permeation scenarios, ranging from an almost completely sealed membrane with only minimal transport at the compound/lipid boundaries, through a previously described sponge-resembling regime,¹⁸ when trying to build putative permeant channels to water-filled channel transport when stabilizing ceramide-pore-inspired hypothetical constructs through a biasing potential (Figure 5).

First, we tested crystal patches in the membrane. Depending on the simulation setup, we observed either dissolution of the crystal structure aggregate to the lipid-water interface or an impenetrable stable membrane-inserted assembly. The strong inter-molecular hydrogen bonding pattern observed in the crystal structure precluded the formation of any water wires or water clusters within the structure, suggesting that a significant re-arrangement of the compounds is required to enable water permeation. Only residual water leakage at the boundary between the crystal patch and the lipids was observed. The crystal structure is stable and displays low permeability as shown in the plot on the bottom. Fitting a line (Figure 5a, red line) to the central part of the curve leads to an estimated permeability of $4 \cdot 10^6$ $\text{H}_2\text{O}/\text{channel}/\text{s}$ (see supplementary information for details on simulation *crys_gaff*).

Next, we attempted the activation of the crystal patch using extensive enhanced sampling simulations to favor internal rearrangements. None of these simulations gave rise to water

transport. In addition to these membrane simulations, we attempted to soak crystal patches in aqueous solution in order to assess how they might accommodate water. Despite an initial separation of the **H1** layers by ca. 2 Å, the crystal patches resisted to any significant water penetration.

Further molecular simulations allowed us to assess whether the water-free crystal superstructures may adaptively re-assemble within the membrane to accommodate any water translocation. These simulation results taken together suggest that the permeant membrane-embedded channel structures are substantially differing from the crystal arrangement. At some stage the compounds may re-arrange into a sponge like form, possibly a stable channel-like form, to give rise to a steady water flow. Three permeation regimes can be distinguished by fitting green, red and orange lines (Figure 5b), with respectively $3 \cdot 10^7$, $8 \cdot 10^8$ and $1 \cdot 10^7$ $\text{H}_2\text{O}/\text{channel}/\text{s}$ permeability (see supplementary information for details on simulation *crys_cgen*). These regimes account for the crystal-like, the sponge-like and the dissolved form of the channels, which are probably most present at low concentration of channels in the membrane.

These observations may likely echo the experimental measurements, indicating a first occurrence of water permeation at low concentrations, a minimal transport regime, originating in the perturbation of the membrane arrangement induced by the presence of the OH-channels in the lipid membrane. It is important to note that the dynamics of different channel-like forms vary significantly, as our hypothetical constructs require artificial stabilization to remain intact within the membrane.

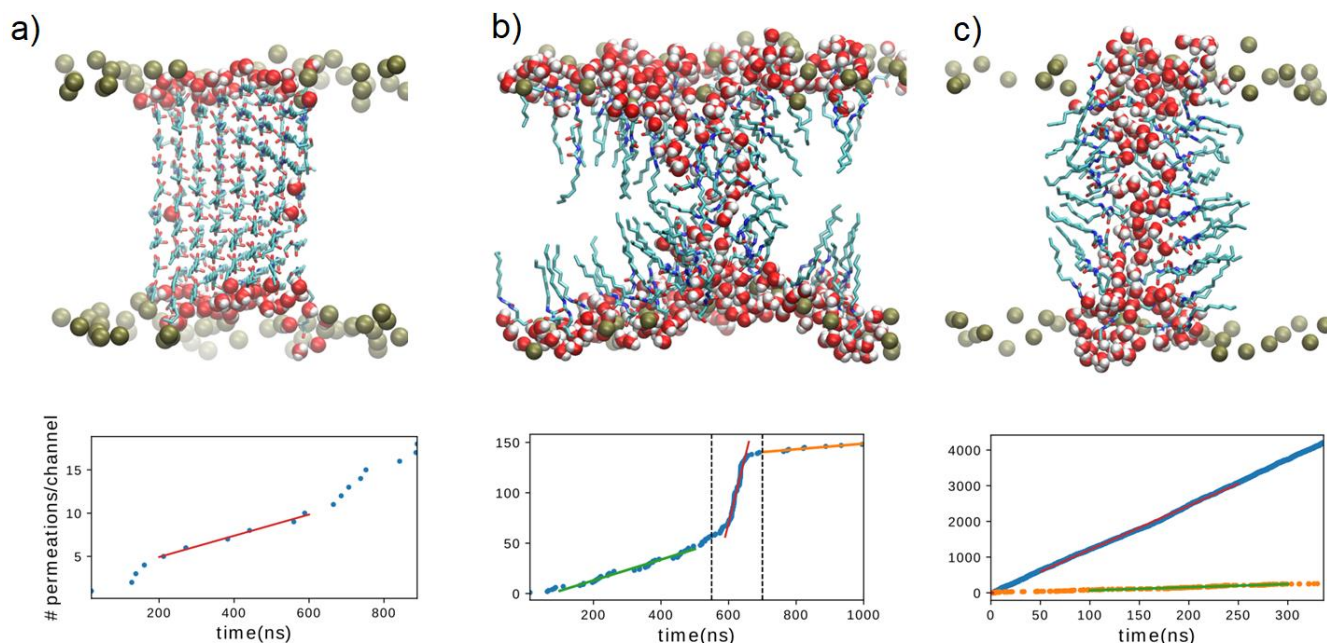


Figure 5. Molecular dynamics simulation results for **H1** hydroxy channels in a fully hydrated lipid bilayer environment. The top row shows snapshots for three transport regimes that were observed: a) a crystal-like patch arrangement after $t = 475$ ns. b) sponge-like structure after $t = 635$ ns c) cylinder-like single membrane-spanning channel in an artificially restrained simulation with a constant force of 1000 kJ/mol (blue, simulation harm1000) and 10 kJ/mol (orange, simulation harm10) at $t = 337$ ns. Channel molecules are drawn in licorice representation. For clarity, lipid and water molecules in the bulk have been omitted. Water molecules within the lipid bilayer or close to the headgroups in contact with H1 are shown in van der Waals representation.

Finally, we constructed a cylindrical pore of **H1** OH-channels resembling the ceramides.²² Our tentative cylindrical channel pore was artificially stabilized in the simulations using harmonic restraints. The corresponding geometry is described in the supplementary information. With such a stabilized cylindrical channel, water transport is indeed observed and can be tuned by increasing the bias from 4×10^8 H₂O/channel/s (weaker force shown in orange in Figure 5c) up to ca. 6×10^9 waters per channel per second (strong force shown in blue in Figure 5c). These observations point to the very subtle balance of interactions that needs to be achieved between the inter-compound and compound-lipid/compound-water interactions to generate high permeable channels as observed at high concentrations of **H1-H6** in the bilayer membrane.

This ensemble of molecular dynamics simulations was further used to check for ion permeation events. Interestingly, in none of the simulations in a bilayer environment did we observe ion permeation during the average 500 ns duration. The limiting case of a simulation with a channel that is forced to stay wide open through strong biasing forces of 1000 kJ/mol lead to some ion permeation. As soon as this restraint force was lowered and the structure gained flexibility, ion permeation stopped.

From a structural point of view, to better understand the membrane modifications operated by this series of compounds, complementary computer simulations are required. These calculations should aim to extend our initial tests with cylindrical pores. The goal is to design and test new constructs of the OH-channel compounds with the challenge to render their architectures sufficiently stable in the membrane environment competing for interactions. So far, we were able to test some artificially stabilized constructs of cylindrical shape, which did feature

permeation. However, by releasing the stabilizing bias, they were torn apart by the favorable lipid interactions in the membrane environment. Our working hypothesis for future tests is to start from the ureic core to assemble the OH-channel compounds, in a similar fashion as has been observed for ceramide-like aggregates of cylindrical shapes²²

Discussion

We show that systematic design of the octyl-ureido-polyols, lead to highly effective and selective Artificial water channels. Simple use of the variable number and sterically different hydroxyl binding sites led to unexpected increased water-transport activities at high concentration of the channels within membranes. They are related to adaptive changes in the structure of the channels, reminiscent with the formation of larger pores for water translocation, showing a 10-fold increase, as observed for natural ceramide transporters.

From a mechanistic point of view, the H-bonding network offered by the hydroxyl OH tetrads can act itself as a recognition-based barrier within highly selective channels for water translocation. The energy penalty for binding either cations or anions would be too high, as either one of the ions can bind either the O or the H at one instance. But on the other hand, an H-bonding water molecule can intervene with both donor and acceptor H-bonds, satisfying the full number of hydrogen bonds within the networks, making possible its facilitated translocation through the gate. The hydroxyl rich cavity of the channels can offer complete amphoteric donor-acceptor hydrogen bonding for the water, while the molecules retain a bulk like bonding environment under confined conditions of the channels. The OH channel architectures are very intriguing systems. It is very

tempting to equate that impermeable crystalline phases or medium permeable sponge like or high permeable cylindrical constructs are stabilized in the bilayer membrane. The unifying assumption is that dynamic H-bonding and hydrophobic interactions would lead to the dominant structures in a membrane environment depending on concentration of the channels and providing pathways for water translocation through the transient water clusters. The water cluster stabilization *via* H-bonding to hydroxyls, confer to these channels self-adaptive behaviors, that make the intrinsic features of water sponges or clusters, particularly self-protective against ions. Having all H bonds saturated via the formation of water wires or water clusters, the cation/anion insertion will impede and interrupt the communication between adjacent water and OH-channels/ pores, explaining the very low ion permeation events. I-quartet single water wires¹⁷ or sponge channels,¹⁸ PAP4 channels¹² or porous organic cages (POCs)¹⁴ are featuring this rather unique enhanced permeability of water clustering conduction systems. This strategy recently described by us, predicts almost two order of magnitude higher permeability of water clusters than of water wires channels.¹⁸

Conclusions

We have demonstrated here that self-assembled adaptive OH-channels can be obtained from simple octyl-ureido-polyol building blocks that present enhanced selective water permeability with almost total ion-rejection across a lipid bilayer. Slight modifications to the number and disposition of hydroxyl-OH functional groups surrounding the hydrophilic cavity of the channels or the increase of the concentration of the compounds in the membrane can drastically change the water permeability. Even with the highest margins of errors, these channels are capable of transporting water at a rate that is one order of magnitude below the natural AQP. More importantly, salt rejection and the extremely low proton permeability of the OH-channels, offer the closest functional mimic of the natural AQP through a synthetic scaffold (Figure 6).

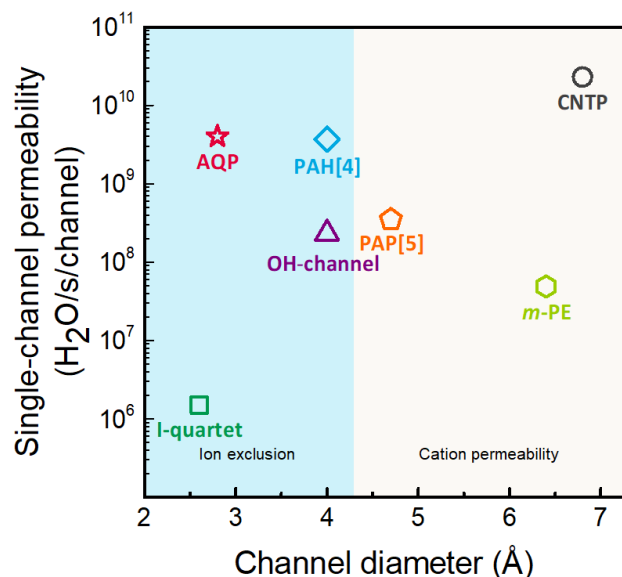


Figure 6. Comparison of single-channel permeabilities vs diameter of AQPs, CNTP and artificial water channels (OH-channel (represented as mean value 4 ± 2.5 Å), I-quartet, PAH4, PAP5, *m*-PE).

The present studies are of tremendous importance showing adaptive behaviors of waters inside the channels depending on the dimension of channels determining the water-conduction behaviors as a function of concentration of OH-channels within membranes.

It is very well known that binding energies define the flux of solutes through an artificial channel. Multiple interactions of molecules with the walls of a channel will impede permeation and result in a reduced flow rate. It was demonstrated that high water permeability values of CNTP, PAP5, PAP4, POCs are attributed to the no-slip translocation mechanism, which is in direct correlation with the weakness of interactions of the water molecules with the walls of the nanotube. This phenomenon gives hydrophobic channels an increase in performance from the point of view of mass transfer. On the other hand, the advantage of hydrophilic channels comes from the very interactional nature of the channel cavity. Having a variety of binding moieties to interact with the transported water, specific binding motifs can be designed. This feature can be exploited, making it possible to design permselectivity filters or tailor the transport mechanism of the channels. They can adapt transient water wires and sponges which are ion selective and water permeable. Within this context this discovery of OH-channels reported here opens up new directions and perspectives in AWCs research,¹⁶ towards the construction of selective membranes for desalination channels,²⁸ biomimicking the functions of natural channels.¹

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website

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Notes

The authors declare no competing financial interest.

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Graphical abstract

Networks of $(\text{OH}\cdots\text{O})_4$ H-bonding quartets for the construction of adaptive water channels

