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Viologen-Cucurbituril Host/Guest Chemistry –Redox Control of Dimerization versus Inclusion

Parastoo Dalvand,^a Katia Nchimi Nono,^b Dinesh Shetty,^c Farah Benyettou,^d Zouhair Asfari,^e Carlos Platas-Iglesias,^f Mark A. Olson,^g Ali Trabolsi^{*d}, and Mourad Elhabiri^{*a}

Two calix[4]arene systems, **C23**⁴⁺ and **C24**⁴⁺ –where 2 corresponds to the number of viologen units and 3-4 corresponds to the number of carbon atoms connecting the viologen units to the macrocyclic core– have been synthesized and led to the formation of [3]pseudorotaxanes when combined with either **CB[7]** or **CB[8]**. The [3]pseudorotaxanes spontaneously dissociate upon reduction of the bipyridinium units as the result of intramolecular dimerization of the two face-to-face viologen radical cations. **CB[7]** and **CB[8]**-based [2]pseudorotaxanes containing monomeric viologen guest models compounds, **MC3**²⁺ and **MC4**⁺, do not undergo decomplexation and dimerization following electrochemical reduction of their bipyridinium units.

Introduction

The stimuli-responsive properties of a molecule with redox-active units can be advantageously used in a broad range of applications including drug delivery and development, electronic displays, electronic memory, batteries, and multi-state switchable materials.¹¹ Among these redox-active molecules, for example metallocenes, tetrathiafulvalene derivatives, naphthalene diimides, and benzoquinones, bipyridinium derivatives¹² (also known as viologens) are of particular interest because of their vibrant electrochromic properties. Bipyridinium dications (**BIPY**²⁺), for example, are typically colorless whereas **BIPY**^{•+} radical monocations¹³ are deep blue in color.¹⁴ Furthermore, these radical cations spontaneously undergo π -dimerization, a process known as pimerization¹⁵ that is driven by radical-radical interactions that favor the formation of the singlet state (*i.e.*, panchromatic absorption)^{14,16} to afford the (**BIPY**^{•+})₂ dimer in aqueous solution. Pimerization of **BIPY**^{•+} radical cations has been used in supramolecular and mechanostereochemical systems to generate translational and rotational molecular movement in many recently reported mechanically interlocked molecules.^{17,19} Furthermore, the rich redox chemistry of viologens also makes them attractive for use in electrochromic, catalytic, and bioanalytical applications.^{20,22}

However, one major complication that limits the use of viologen-based pimerization to develop stimuli-responsive materials is that the system suffers from low dimerization constants (e.g., $\log K_{\text{Dim}} = 2.70$).²³ This results in loosely associated π -dimers in solution, and thus requires the use of high concentrations of the constituent molecular components to drive the binding. The use of low temperatures, micelle environments,²⁴ or macrocyclic hosts with large cavity sizes

such as cucurbiturils (*e.g.*, **CB[8]**), cyclodextrins (*e.g.*, β -CD), or Stoddart's Blue Box^{17,18,25} can help favor the formation of intermolecular π -dimer species. A. E. Kaifer, K. Kim and their coworkers described for the first time the interaction of **MV**²⁺ with cucurbit[n]uril hosts such as **CB[7]** and **CB[8]**.^{26,27,28} With **CB[8]**, it was found that **MV**²⁺ strongly binds in a 1:1 stoichiometry inside the macrocycle cavity, the major driving force being ion-dipole interactions between the positively charged viologen and the oxygen rich host portals atoms. One electron reduction of methyl viologen **MV**²⁺ allowed the formation of a bisradical species ((**MV**^{•+})₂⊂**CB[8]**). The dimerization constant of the **MV**^{•+} in the presence of the **CB[8]** was estimated to be $\sim 2 \times 10^7 \text{ M}^{-1}$ which is about 10⁵ times larger than **MV**^{•+} alone.^{29,28} Supramolecular assemblies based on **CB[8]** has been reviewed recently by A.E. Kaifer, M.D. Garcia et al. recently.³⁰

Alternatively, it has been shown that pre-organizing two or more organic radicals with well-tailored anchoring platforms and appropriate chain lengths can strengthen the intramolecular binding of the π -dimers in solution.³¹ We have previously demonstrated that a hexavalent phosphazene³¹ fulfilled most of the prerequisites including easiness of functionalization, and exhibited efficient and fast pimerization. This compound was applicable towards electrochromic material research and possessed specific recognition properties as a function of the viologen redox state. In addition, alkyl,³²⁻³⁷ porphyrinic,^{38,39} ferrocenyl,⁴⁰ aryl³³⁻⁴⁴, C₆₀,⁴⁵ and calixarene^{46,47} linkers, as well as dendritic^{48,49} and covalent organic polymeric^{50,51} viologen-based materials (Scheme S1 in the ESI) have also been demonstrated to promote intramolecular dimerization of their corresponding bipyridinium radical cations.

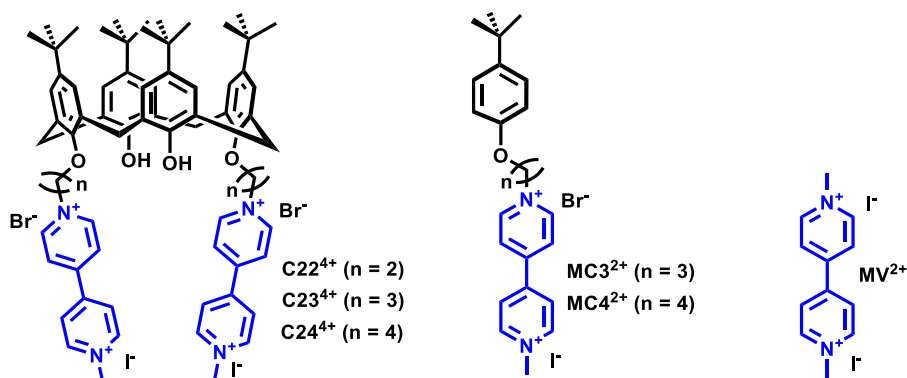


Figure 1. Chemical structures of the calix[4]arene viologen-based systems **C23⁴⁺** and **C24⁴⁺** and the reference compounds **MC3²⁺** and **MC4⁴⁺**. The number 2 stands for the number of viologens while 3-4 correspond to the number of carbons of the chain connecting the terminal viologens to the calix[4]arene backbone. Methyl-viologen (**MV²⁺**) has been used as a model compound in this work. **C22⁴⁺** (PF₆⁻ salt) was not synthesized in this work (see ref. 46) and only consider for DFT calculations.

In line with our ongoing research on viologen-based systems for generating redox-active devices, we turned our attention to calix[4]arene as an inert anchoring platform. Calix[4]arenes can be easily chemically modified and their structural and conformational properties make for an excellent foundation upon which to build increasingly complex architectures. Calix[4]arene can exist in four conformations in solution, the cone, the partial cone, the 1,2-alternate, and the 1,3-alternate. Furthermore, the alkylation selectivity, in regard to the number of substituted phenols, can be controlled by the reaction conditions, thereby giving rise to a broad range of conformers and cavity sizes.⁵²

Herein, we report the synthesis and physicochemical characterization of a series of host-guest monomeric and calix[4]arene-functionalized dimeric viologen-CB-based pseudorotaxane complexes and the structure-property relationships that enable their electrochemical-triggered decomplexation. Calix[4]arene was functionalized with two terminal viologens using spacers of varying length (Figure 1).⁴⁶ The use of synthetic receptors that have both a high affinity and a high selectivity for the binding of guests in water is indeed a very interesting prospect.⁵³⁻⁵⁶ The recognition properties of the viologen-based systems (oxidized and reduced) with **CB[7]** and **CB[8]** were analyzed using a set of complementary analytical methods which include ESI-MS, square wave and cyclic voltammetry, UV-visible-near infrared (NIR) spectrophotometry, ¹H-NMR, and electron paramagnetic resonance (EPR).

Experimental Section

Synthesis

Starting Materials and Solvents. All commercial reagents were purchased from Sigma-Aldrich and used without further purification. 1-methyl-[4,4'-bipyridin]-1-ium iodide (**11**) was prepared according to literature procedures.⁵⁷ 1-(3-bromopropoxy)-4-(tert-butyl)benzene (**12**), 1-(4-bromobutoxy)-4-(tert-butyl)benzene (**13**), 1,3-bis(3-bromopropoxy)-*p*-tert-butyl-calix[4]arene (**15**) and 1,3-bis(4-bromobutoxy)-*p*-tert-

butyl-calix[4]arene (**16**) were also prepared according reported procedures with slight modifications.^{58,59} The dicationic dimethyl viologen **MV**·2I (Figure 1)¹² and *tetrakis-p-tert*-butyl-calix[4]arene **14**^{60,61} were prepared according to literature procedures. Thin layer chromatography (TLC) was used to follow the reactions and was performed on aluminium sheets bearing silica gel 60 F254 (E. Merck). Column chromatography was performed using silica gel (Merck; 40-63 μm). Nuclear magnetic resonance (NMR) spectra were recorded on a Bruker AC 400 with working frequencies of 400 and 100 MHz for ¹H and ¹³C, respectively. Chemical shifts are reported in ppm relative to the signals corresponding to the residual non-deuterated solvents (CDCl₃ (δ = 7.26), CD₃OD (δ = 3.34), and d₆-DMSO (δ = 2.50)). High-resolution electrospray mass (HR-ESI) spectra were measured on a micro Q-TOF (Bruker) spectrometer.

1-(3-(4-(tert-butyl)phenoxy)propyl)-1-methyl-[4,4'-bipyridin]-1,1'-diium bromide iodide (MC3²⁺). 0.8 g (2.95 mmol) of **12** was mixed with 0.88 g (2.95 mmol) of **11** (1-methyl-[4,4'-bipyridin]-1-ium iodide) in 30 mL of acetonitrile (CH₃CN) and the mixture was refluxed for two days. The reaction mixture was filtered and washed with CH₃CN to afford **MC3²⁺** as a red-orange solid (yield: 31%). ¹H NMR (400 MHz, CD₃OD): δ [ppm]: 9.36 [d, 2H, J=6.8 Hz, Ar-H], 9.23 [d, 2H, J=6.8 Hz, Ar-H], 8.70 [t, 4H, J=7.2 Hz, Ar-H], 7.30 [d, 2H, J=8.8 Hz, Ar-H], 6.77 [d, 2H, J=9.2 Hz, Ar-H], 5.04 [t, 2H, J=6.8 Hz, N-CH₂-CH₂], 4.57 [s, 3H, N-CH₃], 4.19 [t, 2H, J=5.6 Hz, O-CH₂-CH₂], 2.66-2.60 [m, 2H, N-CH₂-CH₂-CH₂-O], 1.29 [s, 9H, C(CH₃)₃]. ¹³C NMR (100 MHz, CD₃OD): δ [ppm]: 158.08, 152.66, 152.21, 151.75, 148.88, 148.34, 145.96, 129.04, 128.84, 128.20, 124.39, 115.78, 66.58, 62.13, 35.75, 32.78. Melting point: 258-9 °C. HR-MS for C₂₄H₃₀ON₂: [**MC3²⁺**] calcd. *m/z* = 181.1174 and exp. *m/z* = 181.1172.

1-(4-(4-(tert-butyl) phenoxy)butyl)-1'-methyl-[4,4'-bipyridin]-1,1'-diium bromide iodide (MC4²⁺). 1.58 g (5.53 mmol) of **13** was mixed with 1.5 g (5.02 mmol) of **11** in 30 mL of CH₃CN and the mixture was refluxed for three days. The reaction mixture was filtered and washed with CH₃CN to afford **MC4²⁺** as a yellow-orange solid (yield: 71%). ¹H NMR (400 MHz, CD₃OD): δ [ppm]: 9.36 [d, 2H, J=7.2 Hz, Ar-H], 9.23 [d, 2H, J=6.8 Hz, Ar-H], 8.73 [d, 2H, J=6.8 Hz, Ar-H], 8.70 [d, 2H, J=6.8 Hz, Ar-H], 7.32 [d,

2H, $J=8.8$ Ar-H], 6.88 [d, 2H, $J=8.8$, Ar-H], 4.90 [t, 2H, $J=7.6$ Hz, N-CH₂-CH₂], 4.57 [s, 3H, N-CH₃], 4.10 [t, 2H, $J=6$, CH₂-O-Ar], 2.39-2.32 [m, 2H, N-CH₂-CH₂-CH₂], 2.00-1.93 [m, 2H, O-CH₂-CH₂-CH₂], 1.31 [s, 9H, C(CH₃)₃]. ¹³C NMR (100 MHz, CD₃OD): δ [ppm]: 158.73, 152.13, 148.64, 148.01, 145.52, 129.21, 128.84, 128.09, 115.94, 69.03, 63.95, 35.70, 32.80, 30.55, 27.92. Melting point: 254-5 °C. HR-MS for C₂₅H₃₂ON₂: [MC₄²⁺] calcd. m/z = 188.1252 and exp. m/z = 188.1245.

1,3-bis(3-(1'-methyl-[4,4'-bipyridin]-1,1'-diium)propoxy)-*p*-tert-butyl-calix[4]-arene dibromide diiodide (C23⁴⁺). 1.192 g (4 mmol) of **I1** was added to the flask containing 0.365 g (0.4 mmol) of **I5** and 30 mL of CH₃CN. The mixture was left under reflux for 3 days. The solvent was fully removed under vacuum. The product (**C23⁴⁺**, red-orange solid, yield: 85%) was washed with water, filtrated, and dried under vacuum. ¹H NMR (400 MHz, CD₃OD): δ [ppm]: 9.60 [d, 4H, $J=6.8$ Hz, Ar-H], 9.22 [d, 4H, $J=6.8$ Hz, Ar-H], 8.82 [d, 4H, $J=5.6$ Hz, Ar-H], 8.74 [d, 4H, $J=6.8$ Hz, Ar-H], 7.22 [s, 4H, Ar(calix)-H], 7.07 [s, 4H, Ar(calix)-H], 5.30 [t, 4H, $J=7.2$ Hz, N-CH₂-CH₂], 4.56 [s, 6H, N-CH₃], 4.32-4.23 [d, 8H, Ar-CH₂-Ar and CH₂-O-Ar], 3.50 [d, 4H, $J=13.6$, Ar-CH₂-Ar], 3.02-2.95 [m, 4H, CH₂-CH₂-CH₂], 1.30 [s, 18H, C(CH₃)₃], 1.02 [s, 18H, C(CH₃)₃]. ¹³C NMR (100 MHz, *d*₆-DMSO): δ [ppm]: 150.65, 150.36, 149.54, 148.80, 148.26, 147.52, 146.77, 142.61, 133.91, 128.37, 127.70, 127.00, 126.662, 126.31, 73.26, 58.88, 49.04, 34.92, 37.55, 32.39, 32.29, 31.76, 31.61. Melting point: 295-6 °C. HR-MS for C₇₂H₈₈O₄N₄: [C23⁴⁺+2e-]²⁺ calcd. m/z = 536.3397 and exp. m/z = 536.3382.

1,3-bis(3-(1'-methyl-[4,4'-bipyridin]-1,1'-diium)butoxy)-*p*-tert-butyl-calix[4]arene dibromide diiodide (C24⁴⁺). 0.368 g (0.4 mmol) of **I6** was dissolved in 30 mL of CH₃CN. 1.192 g (4 mmol) of **I1** was added. The mixture was left under reflux for 3 days. The solvent was fully removed under vacuum. The product obtained as an orange solid (**C24⁴⁺**, yield: 67%) was washed with water (due to its poor solubility in water), filtrated and dried under vacuum. ¹H NMR (400 MHz, CD₃OD): δ [ppm]: 9.54 [d, 4H, $J=6.4$ Hz, Ar-H], 9.21 [d, 4H, $J=6.4$ Hz, Ar-H], 8.78 [d, 4H, $J=6.4$ Hz, Ar-H], 8.71 [d, 4H, $J=6.8$ Hz, Ar-H], 7.19 [s, 4H, Ar(calix)-H], 7.05 [s, 4H, Ar(calix)-H], 5.15 [t, 4H, $J=7.6$ Hz, N-CH₂-CH₂], 4.56 [s, 6H, N-CH₃], 4.32 [d, 4H, $J=12.8$, Ar-CH₂-Ar], 4.16 [t, 4H, $J=6.8$ Hz, CH₂-O-Ar], 3.48 [d, 4H, $J=12.8$, Ar-CH₂-Ar], 2.66-2.58 [m, 4H, N-CH₂-CH₂-], 2.25-2.17 [m, 4H, -CH₂-CH₂-O], 1.29 [s, 18H, C(CH₃)₃], 1.02 [s, 18H, C(CH₃)₃]. ¹³C NMR (100 MHz, DMSO): δ [ppm]: 150.79, 150.46, 149.51, 148.89, 148.10, 147.48, 146.74, 142.42, 134.02, 128.41, 127.57, 127.01, 126.56, 126.21, 76.01, 61.22, 49.02, 34.92, 34.52, 32.29, 31.79. Melting point: 224-5 °C. HR-MS for C₇₄H₉₂O₄N₄: [C24⁴⁺+2e-]²⁺ calcd. m/z = 550.3554 and exp. m/z = 550.3561.

Physico-chemistry. The experimental details and related figures are given in the ESI.

DFT calculations. The singlet states of the **C22²⁽⁺⁾**, **C23²⁽⁺⁾** and **C24²⁽⁺⁾** systems were investigated using DFT calculations within the hybrid meta-GGA approximation with the M06-2X functional⁶² and the Gaussian 16 package.⁶³ The standard 6-311G(d,p) basis set was used throughout. Solvent effects (water) were incorporated with the polarized continuum model

(PCM) using default settings and the integral equation formalism.⁶⁴ Frequency calculations were used to confirm that the optimized geometries corresponded to local energy minima on the potential energy surface. The M06-2X/6-311G(d,p) approach in combination with the PCM was found to provide good results for the description of π -stacking interactions of viologen derivatives.⁶⁵

EPR Spectroscopy. EPR spectra were recorded on EPR spectrometer CMS 8400 and measurements were carried out with solutions of the calixarene-*bis*-viologens compounds **C23⁴⁺** and **C24⁴⁺**, along with their model ligands, namely **MC3²⁺** and **MC4²⁺**, in the absence and in the presence of **CB[7]**. This study was performed to evaluate whether or not the radical cations were dimerized or segregated⁶⁶ depending on the composition of the mixture. The radical cations species were produced chemically by addition of an excess of a reducing agent such as sodium dithionite (Na₂S₂O₄). Each solution was prepared at a concentration of $\sim 10^{-4}$ M. The EPR data are available in Figures S17-S20 in the ESI.

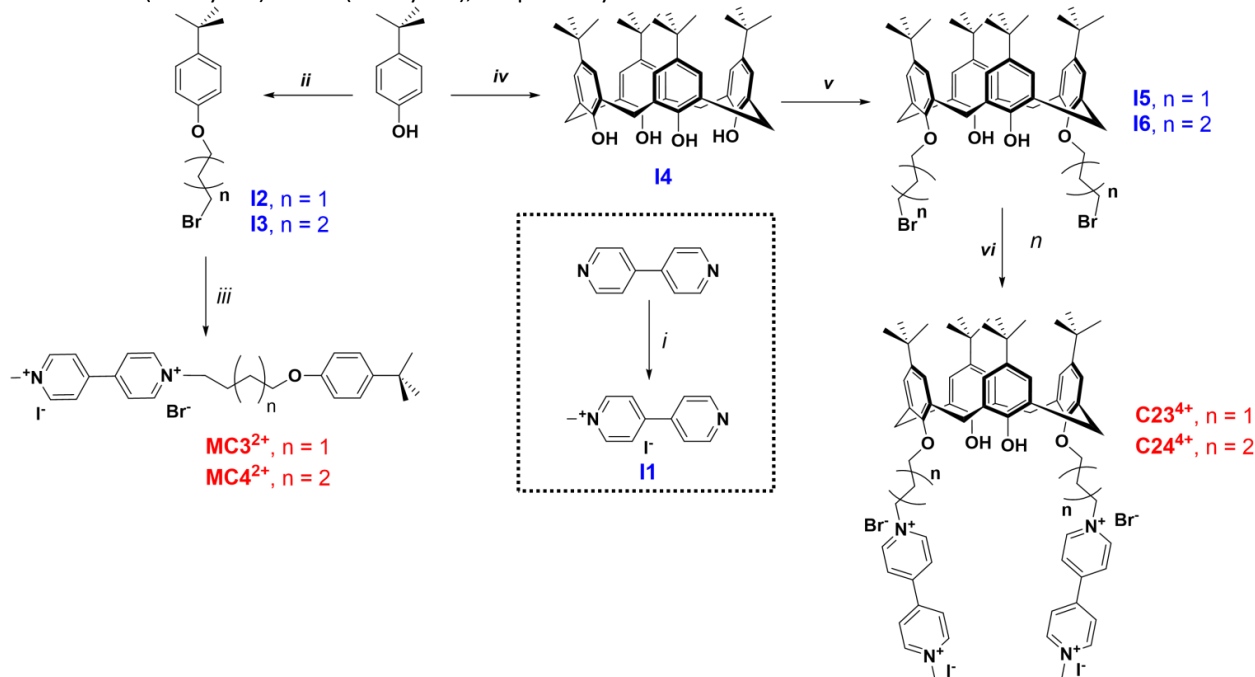
Electrochemistry. All the solutions used for the electrochemical experiments were prepared from 0.1 M TBACl aqueous solution with following concentrations of each species: **MC4²⁺** (C = 6.3 $\times 10^{-5}$ M), **MC3²⁺** (C = 6 $\times 10^{-5}$ M), **C24⁴⁺** (C = 5 $\times 10^{-5}$ M), and **C23⁴⁺** (C = 5 $\times 10^{-5}$ M) with **CB[7]** (C = 2 $\times 10^{-4}$ M). The setup comprises a Gamry Multipurpose instrument (Reference 600) interfaced to a PC. The experiments were performed using a glassy carbon working electrode (0.071 cm², BASi). The electrode surface was polished routinely with 0.05 μ m alumina-water slurry on a felt surface immediately before use. The counter electrode was a Pt coil and the reference electrode was an Ag/AgCl electrode. Cyclic voltammetry (CV) and Square wave differential pulse (SW) were carried out at room temperature in argon-purged H₂O solution between 0 V (initial potential) and -1.2 V (final potential). Chronocoulometry experiments were performed upon stepping from 0 to -0.7 V for 10 s before measuring the response signal corresponding to a radical-cation. Then the potential is stepped from -0.7 V to 0 V allowing measuring the response signal corresponding to the fully oxidized species. The experimental errors on the potential values are ± 10 mV.

Results and Discussion

Synthesis

Two calix[4]arene systems, namely **C23⁴⁺** and **C24⁴⁺**, whereby 2 stands for the number of viologens and 3-4 correspond to the number of carbons of the chain connecting the viologens to the calix[4]arene core, (Scheme 1) were synthesized. For a deeper understanding of the host-guest complexation properties of the viologen derivatives with either **CB[7]** or **CB[8]**, two model compounds, namely **MC3²⁺** and **MC4²⁺**, were also prepared. 4,4'-bipyridine was first monomethylated using 1 equivalent of methyl iodide in acetone leading to **I1** with 63% yield. *Tert*-butyl-phenol was reacted either with 1,3-dibromo-propane or 1,4-dibromo-butane in acetone in the presence of 1 equivalent of potassium carbonate to provide intermediates **I2** or **I3**, respectively. The synthesis of **MC3²⁺** and **MC4²⁺** was then achieved following the reaction of **I2** or **I3** with 1-methyl-4,4'-

bipyridinium (**I1**) in CH₃CN. The synthesis of the *tetrakis-p-tert-butyl-calix[4]arene* **I4** (63% yield) was achieved by reaction of *p-tert-butyl-phenol* with excess of formaldehyde in the presence of NaOH. As described above, **I4** was then reacted with 1,3-dibromo-propane and 1,4-dibromo-butane to lead to intermediates **I5** (92% yield) or **I6** (65% yield), respectively.



Scheme 1. Synthetic route to the **MC3**²⁺ and **MC4**²⁺ monomers as well as the calix[4]arene-*bis*-viologen systems **C23**⁴⁺ and **C24**⁴⁺. Conditions: i) CH₃I, acetone, reflux, 24h, 63%; ii) **I2**: 1,3-dibromopropane, K₂CO₃, acetone, reflux, 48h, yield 70%; **I3**: 1,4-dibromobutane, K₂CO₃, acetone, reflux, 72h, yield 27%; iii) **MC3**²⁺: **I1**, CH₃CN, reflux, 72h, 31%; **MC4**²⁺: **I1**, CH₃CN, reflux, 72h, 71%; iv) NaOH, HCHO, 100 °C, then heating at 195-250 °C in diphenyl ether, toluene, xylene, yield 63%; v) **I5**: 1,3-dibromopropane, K₂CO₃, acetone, reflux, 96h, yield 92%; **I6**: 1,4-dibromobutane, K₂CO₃, acetone, reflux, 96h, yield 65%; vi) iii) **C23**⁴⁺: **I1**, CH₃CN, reflux, 72h, 85%; **C24**⁴⁺: **I1**, CH₃CN, reflux, 72h, 67%.

Characterization of the Thread/CB[7] [n]pseudorotaxanes

Recognition of **MC3²⁺ and **MC4**²⁺ by CB[7].** The UV-Vis absorption spectrophotometric binding titration of **MC3**²⁺ (Figures S1-S2 in the ESI) and **MC4**²⁺ (Figures S3-S4 in the ESI) with **CB[7]** and their corresponding Job plots provided evidence that **MC3**²⁺⊂**CB[7]** and **MC4**²⁺⊂**CB[7]** exist as 1:1 stoichiometric host-guest complexes. Both **MC3**²⁺ and **MC4**²⁺ are characterized by two main absorptions at ~ 225 nm and 260 nm (Table 1) that are attributed to the π-π* transitions of the *tert*-butyl-phenyl and **BIPY**²⁺ chromophoric units, respectively. It was observed that complexation with **CB[7]** induces a hypochromic shift of the absorption band attributed to the **BIPY**²⁺ transitions, while those of the *tert*-butyl-phenyl units remain almost unaltered. This suggested that **CB[7]** is bound to the **BIPY**²⁺ subunit rather than the apparently more hydrophobic alkoxy-*tert*-butyl-phenyl group. This binding behavior is most likely explained by the weaker affinity⁶⁷ of the alkoxy-*tert*-butyl-phenyl moiety for **CB[7]** and steric hindrance of its bulky *tert*-butyl group. The calculated binding constants (log $K_{\text{MC3}^{2+}\text{-CB[7]}} = 4.51(6)$ and log $K_{\text{MC4}^{2+}\text{-CB[7]}} = 4.68(5)$ determined in 0.1 M phosphate buffer, pH 7.0) were found to be lower than that determined for **MV**²⁺, most likely as a consequence of statistical effects. In this case, the two pyridinium aromatic faces of **MV**²⁺ are not sterically blocked to **CB[7]**, while one side of the bipyridinium units of

Reacting these two calix[4]arene intermediates with **I1** then leads to the targeted *bis*-viologen systems **C23**⁴⁺ (85% yield) and **C24**⁴⁺ (67% yield). All these viologen derivatives were isolated in the form of mixed iodide/bromide salts.

MC3²⁺ and **MC4**²⁺ are capped with a *tert*-butyl-phenyl bulky stopper. Consequently, the $K_{\text{MC3}^{2+}}/K_{\text{MV}^{2+}}$ (0.16) and $K_{\text{MC4}^{2+}}/K_{\text{MV}^{2+}}$ (0.24) ratios are close to the statistical value of 0.25.⁶⁸ At the higher concentrations that were used for ¹H NMR (Figures S12 and S13 in the ESI) and the higher **CB[7]/MC3**²⁺ or **CB[7]/MC4**²⁺ host-guest ratios employed for ESI-MS measurements (Figures S8 and S9 in the ESI), other host-guest complexes were observed, namely the [3]pseudorotaxanes, **MC3**²⁺⊂(**CB[7]**)₂ and **MC4**²⁺⊂(**CB[7]**)₂. For these complexes, the ¹H NMR titrations suggested that the second **CB[7]** macrocycle is sitting in close proximity to the *tert*-butyl-phenyl unit (Figure 2a). Binding of the first **CB[7]** indeed mainly affects the β and β' protons of the **BIPY**²⁺ unit, while recognition of the second **CB[7]** macrocycle influences the protons of the *tert*-butyl-phenyl residue (Figures S12 and S13 in the ESI). The inability to quantify the second recognition event leading to the formation of **MC3**²⁺⊂(**CB[7]**)₂ or **MC4**²⁺⊂(**CB[7]**)₂ by absorption spectrophotometry likely resulted from weak spectral variations and their much lower stability constants.

Table 1. Thermodynamic and spectroscopic parameters of [n]pseudorotaxanes formed with **CB[7]**.

Viologen V	log $K_{V-CB[7]}$	V	V \subset CB[7]
		λ (ϵ) nm (10^4 M $^{-1}$ cm $^{-1}$)	λ (ϵ) ϵ^d (10^4 M $^{-1}$ cm $^{-1}$)
MV$^{2+}$	5.30 (2)	227 (2.96) / 257 (2.06)	226 (3.13) / 281 (sh)
MC3$^{2+}$	4.51 (6)	224 (2.29) / 259 (2.17)	223 (2.49) / 255 (1.51)
MC4$^{2+}$	4.68 (5)	224 (2.17) / 260 (2.29)	224 (2.48) / 252 (1.40)
C23$^{4+}$	4.5 (1)	224 (7.42) / 262 (3.20)	225 (7.50) / 275 (2.10)
C24$^{4+}$	4.44 (8)	220 (7.16) / 263 (2.99)	220 (7.29) / 280 (1.83)

Solvent: water buffered at pH 7.0 with 0.1 M Na₂H₂PO₄/NaH₂PO₄; $l = 1$ cm; $T = 25.0(1)$ °C. The error (indicated in brackets) on the stability constants correspond to 3σ with $\sigma =$ standard deviation. sh = shoulder.

To gain further insight into the molecular recognition properties, diffusion coefficients were determined by ¹H NMR-DOSY for **MC3 $^{2+}$** and **MC4 $^{2+}$** both in the absence and in the presence of increasing amounts of **CB[7]** (Table 2). Firstly, a comparison of the diffusion coefficients of the free **MC3 $^{2+}$** and **MC4 $^{2+}$** revealed similar values, which indicated that both model systems share common structural properties (*i.e.*, comparable volume with 303.112 cm³ mol⁻¹ for **MC3 $^{2+}$** and 317.541 cm³ mol⁻¹ for **MC4 $^{2+}$**) notwithstanding the slightly elongated carbon spacer for **MC4 $^{2+}$** . Secondly, the values of the diffusion coefficients experienced a stepwise decrease as a function of the number of equivalents of **CB[7]**, demonstrating the

successive formation of host-guest species of increasing volume and weight (e.g., volume with 927.644 cm³ mol⁻¹ for **MC3 $^{2+}$ ⊂CB[7]**). Lastly, the diffusion coefficients measured for **MC3 $^{2+}$** and **MC4 $^{2+}$** using the same amount of **CB[7]** are comparable to one another, which suggested that the [2]pseudorotaxanes, **MC3 $^{2+}$ ⊂CB[7]** and **MC4 $^{2+}$ ⊂CB[7]**, and the [3]pseudorotaxanes, **MC3 $^{2+}$ ⊂(CB[7])₂** and **MC4 $^{2+}$ ⊂(CB[7])₂**, are roughly structurally equivalent. The inability to determine the stability constant of [3]pseudorotaxane prevents the evaluation of the diffusion coefficients specific to each inclusion complex (*i.e.*, [2]- and [3]pseudorotaxanes).

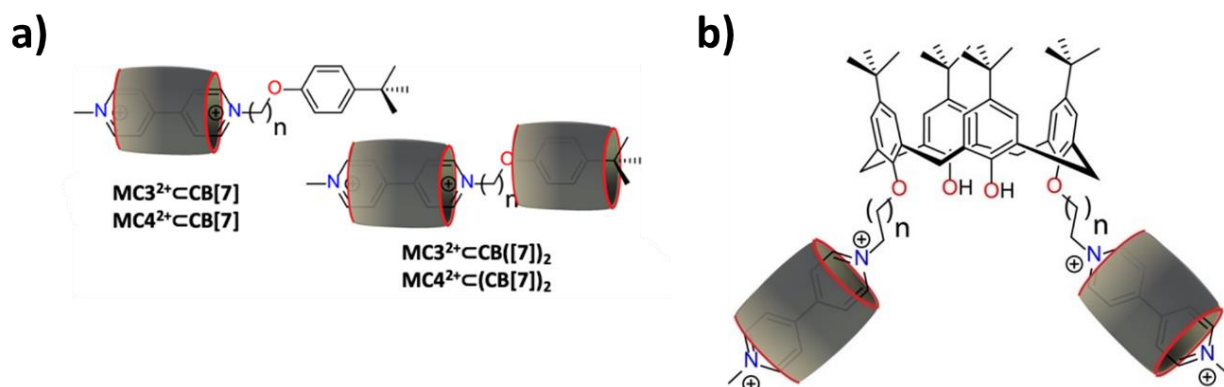


Figure 2. Schematic representation of a) the [2]pseudorotaxanes **MC3 $^{2+}$ ⊂CB[7]** and **MC4 $^{2+}$ ⊂CB[7]** (**CB[7]** mainly resides on the bipyridinium group) and of **MC3 $^{2+}$ ⊂(CB[7])₂** and **MC4 $^{2+}$ ⊂(CB[7])₂** (*i.e.*, the hydrophobic cavity of the second **CB[7]** is sitting close to the terminal phenyl group); and b) the [3]pseudorotaxanes **C23 $^{4+}$ ⊂(CB[7])₂** and **C24 $^{4+}$ ⊂(CB[7])₂** (*i.e.*, the **CB[7]** mainly resides on the bipyridinium groups). $n = 2$, or 3. **CB[7]** is represented with the grey cylinder

Recognition of C23 $^{4+}$ and C24 $^{4+}$ by CB[7]. Altogether, absorption binding titrations (Figures S6 and S7 in the ESI), ¹H-NMR data (Figures S14 and S15 in the ESI), and ESI-MS experiments (Figures S10 and S11 in the ESI) demonstrated that the [3]pseudorotaxanes, **C23 $^{4+}$ ⊂(CB[7])₂** and **C24 $^{4+}$ ⊂(CB[7])₂**, formed predominantly. Hypochromic and bathochromic shifts of the absorption bands which correspond to the π - π^* transitions of the **BIPY $^{2+}$** cores were observed in both cases (Table 1) as a consequence of the encapsulation of the terminal **BIPY $^{2+}$** chromophores within the hydrophobic macrocyclic cavity of **CB[7]**. The absorption bands which correspond to the π - π^* transitions for the calix[4]arene core remained unaffected by **CB[7]** complexation. As previously shown with the model systems, **MC3 $^{2+}$** and **MC4 $^{2+}$** , the presence of a hydrophobic

alkoxy substituent did not alter the molecular recognition properties. As a result, the two **CB[7]** macrocycles remained bound to the middle of each of the two **BIPY $^{2+}$** terminal electrophores as evidenced by ¹H NMR spectroscopy (Figures S14 and S15 in the ESI). The apparent stability constants ($\log K^*_{CB23^{4+}\subset(CB[7])_2} = 4.5(1)$ and $\log K^*_{CB24^{4+}\subset(CB[7])_2} = 4.44(8)$) were found to be close to those evaluated for the corresponding models, **MC3 $^{2+}$** and **MC4 $^{2+}$** (Table 1). Again, these stability constant values were found to be lower than those measured for **MV $^{2+}$** ($\log K_{MV^{2+}\subset CB[7]} = 5.30(2)$) as a consequence of the presence of the bulky calix[4]arene core (Figure 2b).

Diffusion coefficients were also determined for **C23 $^{4+}$** and **C24 $^{4+}$** both in the absence and the presence of **CB[7]**. In contrast to the values measured for the model ligands, the D value

measured for **C23⁴⁺** was found to be lower than that of **C24⁴⁺**. One would expect that the extended spacer of **C24⁴⁺** should increase the global volume of the molecule (*i.e.*, calculated volumes of 824.390 cm³ mol⁻¹ for **C23⁴⁺** and 872.891 cm³ mol⁻¹ for **C24⁴⁺**), thereby causing it to diffuse slower in solution. This unexpected result thus might be explained by the slightly more flexible arms of **C24⁴⁺** which should minimize the steric interactions of the viologen units with the bulky calix[4]arene

core, thereby resulting in markedly different solvation properties with respect to **C23⁴⁺**. As a consequence, the diffusion coefficient decreases in the presence of **CB[7]** for **C24⁴⁺**, while it increases for **C23⁴⁺**, and both converge to a comparable value, the volumes of host-guest species with **C23⁴⁺** and **C24⁴⁺** being then globally equivalent in the presence of **CB[7]**.

Table 2. Diffusion coefficients of **MC3²⁺**, **MC4²⁺**, **C23⁴⁺** and **C24⁴⁺** in the absence and the presence of **CB[7]**.^a

$D \text{ (} \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}\text{)}$									
No CB[7]		+ 1 eq CB[7]		+ 2 eq CB[7]		No CB[7]		+ 2 eq CB[7]	
MC3²⁺	MC4²⁺	MC3²⁺	MC4²⁺	MC3²⁺	MC4²⁺	C23⁴⁺	C24⁴⁺	C23⁴⁺	C24⁴⁺
6.51	6.57	4.89	4.17	3.56	3.29	2.58	4.30	3.85	3.11

^a Measured by DOSY experiments on a 600 MHz spectrometer, D₂O, 298 K. The errors on these values are estimated to be less than 5%.

Characterization of the Radical Cations

Intermolecular versus Intramolecular Dimerization of the Radical Cations. We have demonstrated that **MC3^{•+}** and **MC4^{•+}** rapidly pimerize (Figures S21 and S22 in the ESI) in aqueous solution with a log K_{Dim} value of 3.4. This value is very similar to the value previously determined for the benzyl methyl viologen radical cation **BMV^{•+}** (log $K_{\text{Dim}} = 3.46(5)$),³¹ but much higher than that reported for the methyl viologen radical cation **MV^{•+}** (log $K_{\text{Dim}} \sim 2.5\text{--}2.9$ ^{69,70}). The values of K_{Dim} are also similar to those measured for viologens decorated with hydrophobic alkyl substituents.¹⁵ This suggested that increasing the hydrophobicity around the **BIPY^{•+}** unit with aryl/alkyl substituents favors the intermolecular pimerization in aqueous solution. Moreover, we could again demonstrate that the extended spacer for **MC4^{•+}** does not drastically alter its pimerization since similar K_{Dim} values were calculated for **MC3^{•+}** and **MC4^{•+}**. These two radical cations are characterized by intense and structured absorption bands in the visible region (**MC3^{•+}**: $\lambda_{\text{max}} \sim 602 \text{ nm}$, $\epsilon^{602} = 1.01 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$; **MC4^{•+}**: $\lambda_{\text{max}} \sim 602 \text{ nm}$, $\epsilon^{602} = 9.56 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$, Figures S21 and S22 in the ESI), in agreement with the spectroscopic parameters determined for **MV^{•+}** (600–606 nm)¹⁵, **NOTEREF_Ref75436386** \h 2370,71 and **BMV^{•+}** (600 nm).³¹ Formation of the radical cation dimers induced a significant hypsochromic shift of the absorption band at $\sim 600 \text{ nm}$ ($\Delta\lambda \sim 50 \text{ nm}$) and gave rise to intense absorption bands corresponding to radical–radical transitions in the NIR region ($\lambda_{\text{max}} > 850\text{--}900 \text{ nm}$).¹⁵ These UV-Vis absorption studies also demonstrated that no significant pimerization occurs unless the concentration of the radical cations is high enough ($> 10^{-3} \text{ M}$, Figures S21 and S22 in the ESI). As previously reported,^{32–46} preorganization of the **BIPY²⁺** electrophores to within close proximity of one another around a robust and inert molecular platform can enhance their pimerization, leading to very stable π -dimers in aqueous solution.

Upon chemical generation of the radical cations, **C23^{2(•+)}** and **C24^{2(•+)}** (Figures S27 and S28 in the ESI), intense absorption bands appeared at $\sim 530 \text{ nm}$ (Figure 3). An additional intense absorption band centered at 1072 nm for **C23^{2(•+)}** and at 925 nm for **C24^{2(•+)}** were observed and are unambiguously assigned to the intramolecular charge resonance that occurs in the dimerized viologen radical cation species. This feature indicated that a two-electron reduction of **C23⁴⁺** and **C24⁴⁺** leads to two

radical divalent cations, **C23^{2(•+)}** and **C24^{2(•+)}**, that spontaneously pimerize intramolecularly, leading to a stable radical dimeric species.

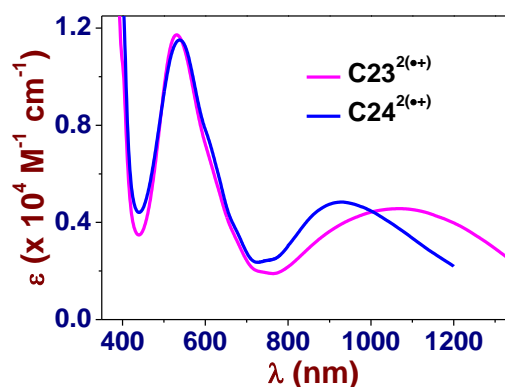


Figure 3. Electronic UV-Vis-NIR absorption spectra of **C23^{2(•+)}** and **C24^{2(•+)}** in water. Solvent: water buffered at pH 7.0 with 0.1 M Na₂HPO₄/NaH₂PO₄, $T = 25.0(1) \text{ } ^\circ\text{C}$.

The NIR absorption spectroscopic signatures of the π -dimerized complexes were demonstrated to be closely related to the extent of overlapping between both π -orbitals of the **BIPY^{•+}** radicals.⁴⁶ When a propyloxy chain links the viologens to the calix[4]arene moiety (**C23^{2(•+)}**: $\lambda_{\text{max}} \sim 960 \text{ nm}$ in CH₃CN⁴⁶ and $\lambda_{\text{max}} \sim 1070 \text{ nm}$ in water), DFT calculations at the BLYP-D3/DZVP level suggested that only one pyridinium ring per viologen radical is involved in the intramolecular π -dimerization process. This result, which is imposed by geometric constraints that is inherent to the length and structure of the propyl chain, is not observed with the ethyloxy-derived calix[4]arene-*bis*-viologen (**C22^{2(•+)}**: $\lambda_{\text{max}} \sim 860 \text{ nm}$ in CH₃CN⁴⁶), which affords a fully face-to-face stacked arrangement. DFT calculations performed at the M06-2X/6-311G(d,p) level (Figure 4) provided a minimum energy conformation for **C22^{2(•+)}** in which the two **BIPY^{•+}** units display an almost perfect face-to-face interaction. Similarly, the **BIPY^{•+}** units in **C23^{2(•+)}** are twisted with respect to each other by an angle of $\sim 78^\circ$, reducing the overlap between the corresponding π -orbitals. Increasing the length from 3 to 4 carbons (butyloxy derivative, **C24^{2(•+)}**: $\lambda_{\text{max}} \sim 925 \text{ nm}$ in water) does not completely reinstate the face-to-face dimerized π -stacking orientation, but it does increase the overlap of the π -orbitals with respect to **C23^{2(•+)}** with a twist angle of $\sim 60^\circ$. Thus,

the nature and size of the linker between the calix[4]arene platform and the terminal **BIPY**^{•+} radical cations, as well as the nature of the platform, with the calix[4]arene displaying a cone conformation, appear to be crucial factors that govern the

arrangement of the π -dimers. It is noteworthy that preorganization of the designed systems apparently allows for efficient intramolecular dimerization of the two viologens upon reduction in a range of solvents (*e.g.*, water and CH₃CN⁴⁶).

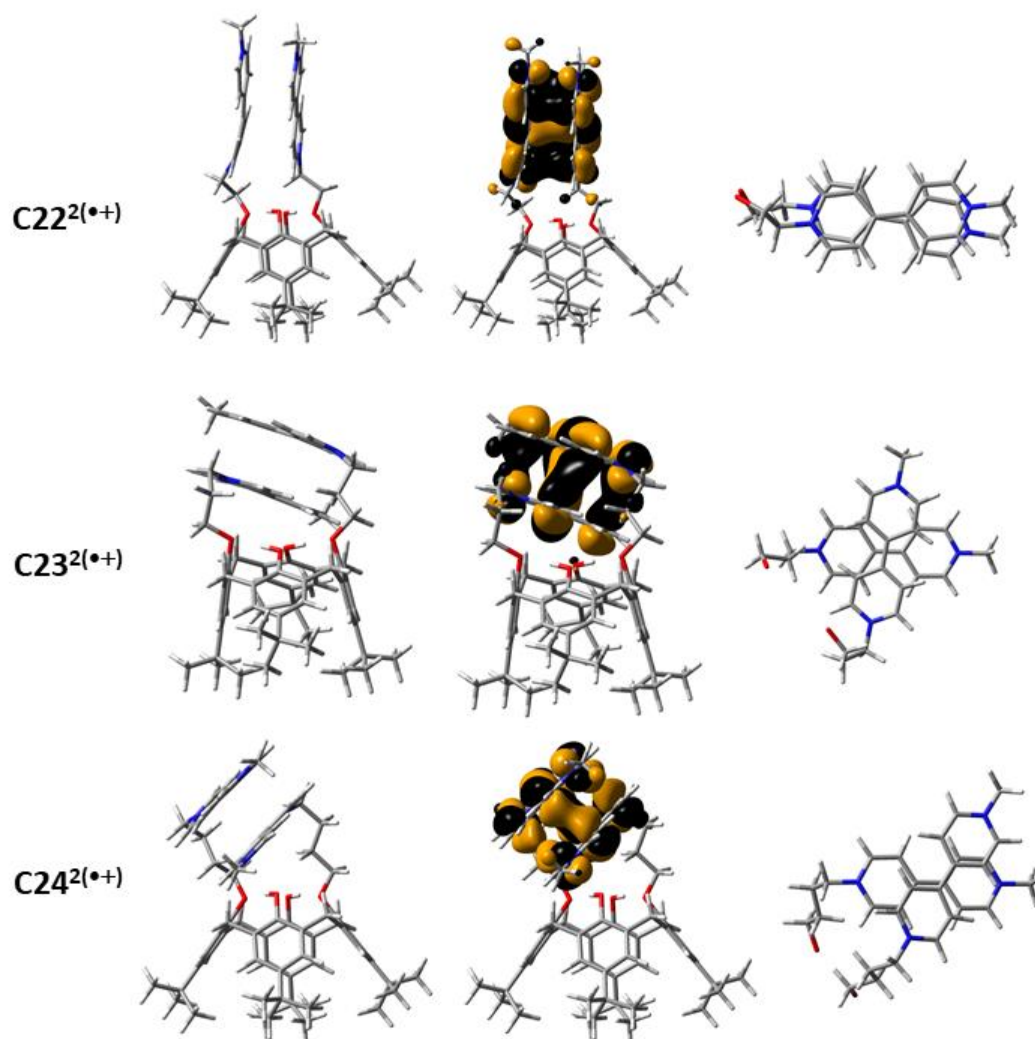


Figure 4. Geometries obtained with DFT (M062X/6-311G(d,p)) (left panel), the corresponding HOMOs (central panel) and detail of the arrangements of the redox-active **BIPY**^{•+} units within the two electrons reduced calix[4]arene-bis-viologens **C22**^{2(•+)}, **C23**^{2(•+)} and **C24**^{2(•+)} that undergo intramolecular pimerization.

Reduction of the [n]Pseudorotaxanes with **CB[7]**

Monocationic Monoradicals MC3^{•+} and MC4^{•+}. Cyclic voltammetry (Figure S16 in the ESI), square-wave voltammetry (Figure 5), and EPR (Figures S17 and S18 in the ESI) measurements in phosphate-buffered solutions at pH 7 were performed on the model systems, **MC3**²⁺ and **MC4**²⁺, in the absence and presence of **CB[7]**. Both systems were

characterized by two successive one-electron reversible redox waves: $E_{1/2}^{1/2}(\text{MC3}^{2+} \rightarrow \text{MC3}^{\bullet+}) = -0.56 \text{ V}$ and $E_{1/2}^{2/2}(\text{MC3}^{\bullet+} \rightarrow \text{MC3}^0) = -0.83 \text{ V}$; $E_{1/2}^{1/2}(\text{MC4}^{2+} \rightarrow \text{MC4}^{\bullet+}) = -0.56 \text{ V}$ and $E_{1/2}^{2/2}(\text{MC4}^{\bullet+} \rightarrow \text{MC4}^0) = -0.83 \text{ V}$. In the presence of three equivalents of **CB[7]**, where it is assumed that **MC3**²⁺⊂**CB[7]** and **MC4**²⁺⊂**CB[7]** predominate, both redox waves shifted slightly to more negative potentials (Table 3) while retaining their reversible shapes.

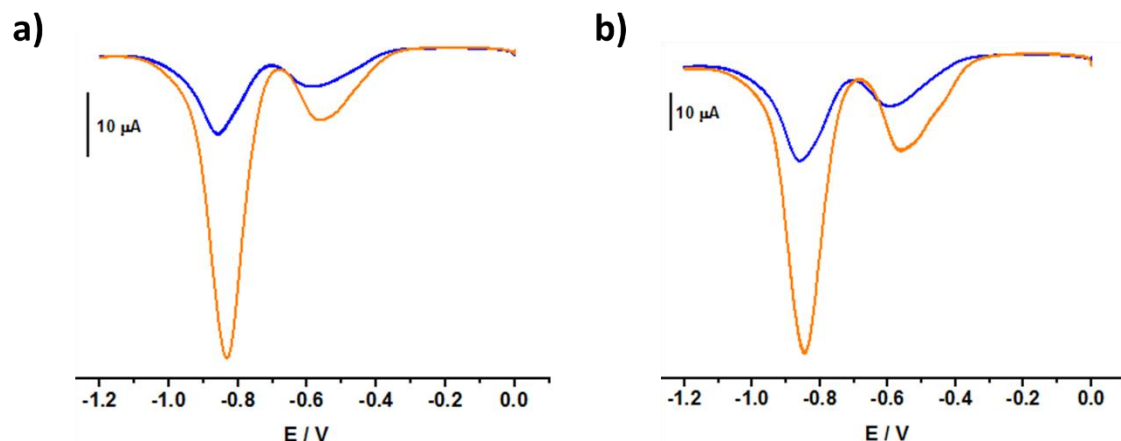


Figure 5. Square Wave voltammograms of a) MC3^{2+} , $[\text{MC3}^{2+}] = 0.05 \text{ mM}$ and b) MC4^{2+} , $[\text{MC4}^{2+}] = 0.06 \text{ mM}$ in the absence (orange) and the presence of **CB[7]** (blue, 3 equivalents of **CB[7]**). All voltammograms were recorded in Argon-purged phosphate buffer solutions (pH 7) at 298 K (E versus Ag/AgCl).

These shifts corresponded to the signatures of the relative affinities of **CB[7]** for the different redox states of the model systems. Assuming a $\log K_{\text{MC3}^{2+}\text{-CB[7]}}$ value of 4.51(6) for $\text{MC3}^{2+}\text{-CB[7]}$ (Figure S1 in the ESI), a $\log K_{\text{MC3}^{0+}\text{-CB[7]}}$ of ~ 4 -4.2 and a $\log K_{\text{MC3}^{0}\text{-CB[7]}}$ of ~ 3.5 -3.9 values were accordingly calculated.^{26,29} Similarly, using a $\log K_{\text{MC4}^{2+}\text{-CB[7]}}$ value of 4.68(5) (Figure S3 in the ESI), a $\log K_{\text{MC4}^{0+}\text{-CB[7]}}$ of ~ 4.2 and a $\log K_{\text{MC4}^{0}\text{-CB[7]}}$ of ~ 3.8 were calculated. Using an electrochemical approach, these calculated values were in reasonably good agreement, within error, with those derived from direct absorption binding titrations ($\log K_{\text{MC3}^{0+}\text{-CB[7]}} = 3.8(4)$ (Figure S24 in the ESI) and $\log K_{\text{MC4}^{0+}\text{-CB[7]}} = 4.54(5)$ (Figure S26 in the ESI). In contrast to BMV^{2+} , where translational motion of the **CB[7]** macrocycle occurred upon electrochemical reduction of the

bipyridinium unit,³¹ the **CB[7]** macrocycle remained firmly bound to the BIPY^{2+} electroactive units of the MC^{4+} derivatives regardless of their oxidation state.

Assuming that only the first one-electron reduction process dominates from 0 V to -0.7 V (Figure S16 in the ESI and Figure 5), the diffusion coefficients of MC3^{2+} and MC4^{2+} in the absence and in the presence of **CB[7]** (Table 4) can be evaluated by chronocoulometric experiments. The differences observed between the values of the diffusion coefficients of $\text{MC3}^{0+}/\text{MC3}^{2+}$ and $\text{MC4}^{0+}/\text{MC4}^{2+}$ could be an indication of the dimerization of the model systems upon the one-electron reduction reactions. Addition of **CB[7]** suppresses the pimerization process in favor of the inclusion complexes. As a consequence, the inclusion complexes with **CB[7]** have lower diffusion coefficients than the corresponding free viologens.

Table 3. Values of $E_{1/2}$ (in V) for MC3^{2+} (0.05 mM), MC4^{2+} (0.063 mM), C23^{4+} (0.05 mM) and C24^{4+} (0.05 mM) measured by cyclic voltammetry (CV) and square wave voltammetry (SWV) in the absence and in the presence of 3 equivalents of **CB[7]** for MC3^{2+} , MC4^{2+} and C23^{4+} , and 4 equivalents of **CB[7]** for C24^{4+} .

Species	$E_{1/2}^1$ (CV)	$E_{1/2}^1$ (SW)	$E_{2/2}^2$ (CV)	$E_{2/2}^2$ (SW)
MC3^{2+}	-0.56 V	-0.56 V	-0.83 V	-0.83 V
$\text{MC3}^{2+} + 3 \text{ eq CB[7]}$	-0.60 V	-0.59 V	-0.85 V	-0.86 V
MC4^{2+}	-0.56 V	-0.56 V	-0.83 V	-0.84 V
$\text{MC4}^{2+} + 3 \text{ eq CB[7]}$	-0.58 V	-0.58 V	-0.83 V	-0.86 V
C23^{4+}	-0.42 V	-0.43 V	-0.84 V	-0.85 V
$\text{C23}^{4+} + 3 \text{ eq CB[7]}$	-0.44 V	-0.46 V	-0.84 V	-0.85 V
C24^{4+}	-0.45 V	-0.47 V	-0.85	-0.85
$\text{C24}^{4+} + 4 \text{ eq CB[7]}$	-0.50	-0.51	-0.85	-0.85

Table 4. Diffusion coefficients D ($\text{cm}^2 \text{ s}^{-1}$) of MC3^{2+} , MC4^{2+} , C23^{4+} and C24^{4+} measured by chronocoulometry in H_2O (0.1 M TBACl) in the absence and the presence of **CB[7]**. The errors on these values are 10%.

No CB[7] / D	+ CB[7] / D
$\text{MC3}^{0+} / 0.7 \times 10^{-5}$	$\text{MC3}^{0+} + 2 \text{ eq CB[7]} / 1.9 \times 10^{-5}$
$\text{MC3}^{2+} / 5.4 \times 10^{-5}$	$\text{MC3}^{2+} + 2 \text{ eq CB[7]} / 14.2 \times 10^{-5}$
$\text{MC4}^{0+} / 3.1 \times 10^{-5}$	$\text{MC4}^{0+} + 2 \text{ eq CB[7]} / 6.2 \times 10^{-5}$

$MC4^{2+} / 81.0 \times 10^{-5}$	$MC4^{2+} + 2 \text{ eq CB[7]} / 22 \times 10^{-5}$
$C23^{2+} / 1.41 \times 10^{-5}$	$C23^{2+} + 3 \text{ eq CB[7]} / 2.9 \times 10^{-5}$
$C23^{2+} / 7.58 \times 10^{-5}$	$C23^{2+} + 3 \text{ eq CB[7]} / 3.91 \times 10^{-5}$
$C24^{2+} / 0.81 \times 10^{-5}$	$C24^{2+} + 4 \text{ eq CB[7]} / 0.16 \times 10^{-5}$
$C24^{2+} / 20.1 \times 10^{-5}$	$C24^{2+} + 4 \text{ eq CB[7]} / 42.1 \times 10^{-5}$
$C23^{2+} / 1.41 \times 10^{-5}$	$C23^{2+} + 3 \text{ eq CB[7]} / 2.9 \times 10^{-5}$
$C23^{2+} / 7.58 \times 10^{-5}$	$C23^{2+} + 3 \text{ eq CB[7]} / 3.91 \times 10^{-5}$
$C24^{2+} / 0.81 \times 10^{-5}$	$C24^{2+} + 4 \text{ eq CB[7]} / 0.16 \times 10^{-5}$
$C24^{2+} / 20.1 \times 10^{-5}$	$C24^{2+} + 4 \text{ eq CB[7]} / 42.1 \times 10^{-5}$

Investigations of the UV-Vis-NIR absorption spectra of $MC3^{2+}$ and $MC4^{2+}$ in water at pH 7.0 in the absence or in the presence of a reducing agent were carried out to evaluate the effect of $CB[7]$ addition (Figures S23 and S25 in the ESI). $MC3^{2+} \subset CB[7]$ and $MC4^{2+} \subset CB[7]$ both clearly exhibited a significant hypochromic shift of the visible absorption band associated

with the radical cation with increasing concentration of $CB[7]$, strong evidence for the inclusion of the $BIPY^{2+}$ radical cation within the hydrophobic cavity of $CB[7]$. This binding event effectively hampers radical cation pimerization.

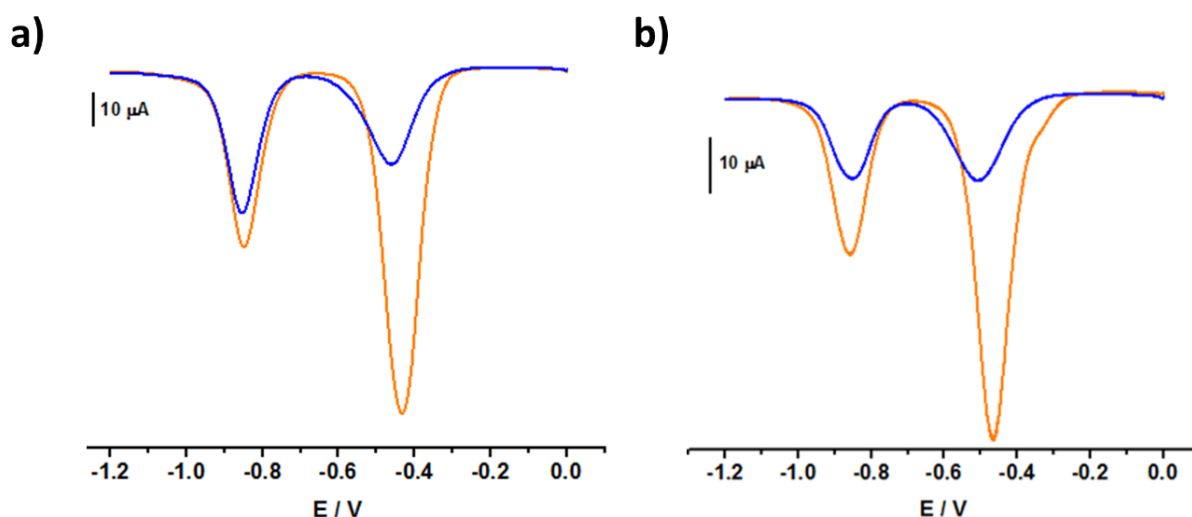


Figure 6. Square Wave voltammograms of a) $C23^{4+}$, $[C23^{4+}] = 0.05 \text{ mM}$ and b) $C24^{4+}$, $[C24^{4+}] = 0.05 \text{ mM}$ in the absence (orange line) and the presence of $CB[7]$ (blue line, respectively 3 and 4 equivalents of $CB[7]$). All voltammograms were recorded in Argon-purged phosphate buffer solutions (pH 7) at 298 K (E versus $Ag/AgCl$).

Bisradicals $C23^{2(•+)}$ and $C24^{2(•+)}$. Cyclic (Figure S16 in the ESI) and square wave (Figure 6) voltammetry studies were undertaken to evaluate the impact of $CB[7]$ addition on the intramolecular pimerization of the $BIPY^{2+}$ radical cations of the calix[4]arene-*bis*-viologens, $C23^{4+}$ and $C24^{4+}$. The electrochemical data clearly showed two distinct and reversible redox waves. The relative amplitudes and shapes of the peaks for each wave at the anode (oxidation) are similar to those observed at the cathode (reduction). Such a pattern is an indication of the redox reversibility of the system.

The square wave voltammograms of $C23^{4+}$ and $C24^{4+}$ (Figure 6) displayed two reversible two-electron reduction processes: $E_{1/2}^1 (C23^{4+} \rightarrow C23^{2(•+)}) = -0.43 \text{ V}$ and $E_{1/2}^2 (C23^{2(•+)} \rightarrow C23^0) = -0.85 \text{ V}$; $E_{1/2}^1 (C24^{4+} \rightarrow C24^{2(•+)}) = -0.47 \text{ V}$ and $E_{1/2}^2 (C24^{2(•+)} \rightarrow C24^0) = -0.85 \text{ V}$. Interestingly, for the [3]pseudorotaxanes, $C23^{4+} \subset (CB[7])_2$ and $C24^{4+} \subset (CB[7])_2$, no shifts were observed for the second reduction wave (Table 3), which would have

corresponded to the following electrochemical processes: $C23^{2(•+)} \subset (CB[7])_2 \rightarrow C23^0 \subset (CB[7])_2$ and $C24^{2(•+)} \subset (CB[7])_2 \rightarrow C24^0 \subset (CB[7])_2$. Rather, our observations indicated that after the first two-electron reduction process, the threads behave as if they are unbound $C23^{2(•+)}$ and $C24^{2(•+)}$. In other words, the first reduction of $C23^{4+} \subset (CB[7])_2$ and $C24^{4+} \subset (CB[7])_2$ induced dethreading of the $CB[7]$ macrocycles, which is spontaneously followed by intramolecular dimerization of the two terminal $BIPY^{2+}$ groups, leading to the formation of intramolecular dimeric species, $C23^{2(•+)}$ and $C24^{2(•+)}$. These results were consistent with the absorption spectrophotometric analyses (Figures S27 and S28 in the ESI) and EPR behavior (Figures S19 and S20 in the ESI) of the $C23^{4+}/CB[7]$ and $C24^{4+}/CB[7]$ [3]pseudorotaxanes.

Chronocoulometry was also used to evaluate the diffusion coefficients of the fully oxidized and radical cationic states of

the different species in aqueous solution. During the two electrons reduction process (0 V \rightarrow -0.7 V) the fully oxidized species predominate and diffuse to the electrode surface where they are reduced. By measuring the reduction rates, the diffusion constants for C23^{4+} and C24^{4+} species can be calculated in the absence and in the presence of $\text{CB}[7]$ (Table 4). Alternatively, by setting the voltage to -0.7 V and performing the oxidation process, the rates of oxidation and the diffusion coefficients of the corresponding radical cationic species can be measured as well. In the absence of $\text{CB}[7]$, larger diffusion coefficients were observed for C23^{4+} and C24^{4+} than for their corresponding fully reduced forms. This could be explained by the ability of higher charged species to diffuse faster towards or away from an electrode. In the presence of $\text{CB}[7]$, larger diffusion coefficients were measured for the [3]pseudorotaxanes $\text{C23}^{4+}\text{-(CB}[7])_2$ and $\text{C24}^{4+}\text{-(CB}[7])_2$ than the corresponding reduced species (*i.e.*, upon reduction, a dethreading of the $\text{CB}[7]$ occurs as a consequence of favored intramolecular pimerization with respect to the recognition process). Here also, one would expect that the [3]pseudorotaxanes $\text{C23}^{4+}\text{-(CB}[7])_2$ and $\text{C24}^{4+}\text{-(CB}[7])_2$ display larger hydrodynamic size by comparison with $\text{C23}^{2(++)}$ or $\text{C24}^{2(++)}$. Markedly different solvation of the latter species is suggested to explain these peculiar properties. The close diffusion coefficients of the radical cation species in the absence and the presence of $\text{CB}[7]$ indicate that the dimerization of the viologen radical cations induce a dethreading of the $\text{CB}[7]$ macrocycles.

Recognition of the Viologen Derivatives by $\text{CB}[8]$

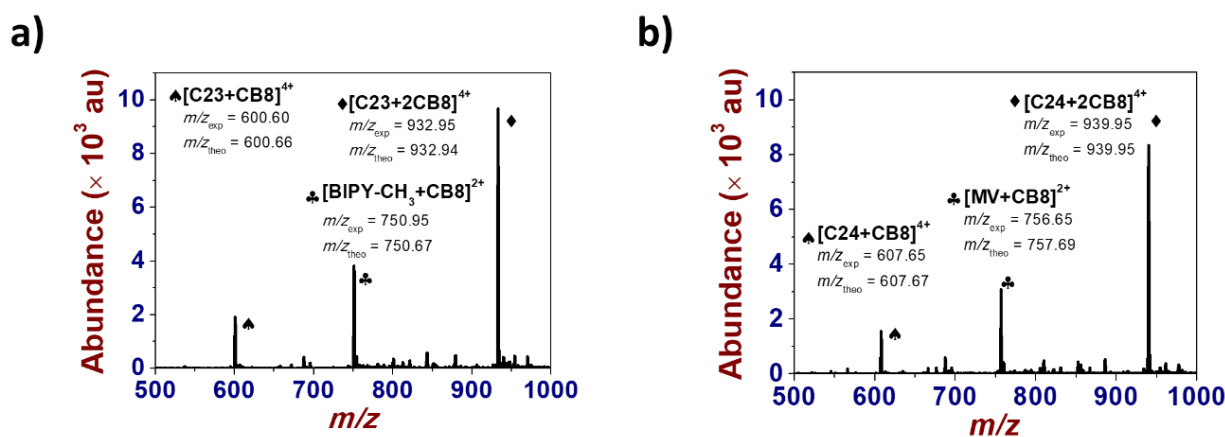


Figure 7. ESI mass spectra of inclusion complexes of a) C23^{4+} and b) C24^{4+} (bottom) with $\text{CB}[8]$. a) $[\text{C23}^{4+}]_0 = 5 \times 10^{-5}$ M; $[\text{CB}[7]]_0 = 10^{-4}$ M. Solvent: H_2O ; positive mode. $V_c = 200$ V. b) $[\text{C24}^{4+}]_0 = 5 \times 10^{-5}$ M; $[\text{CB}[7]]_0 = 10^{-4}$ M. Solvent: H_2O ; positive mode. $V_c = 100$ V.

Conclusions

Two calix[4]arene-*bis*-viologen systems, namely C23^{4+} and C24^{4+} were synthesized. The host-guest properties were extensively studied using a large set of analytical methods and led to the characterization of [3]pseudorotaxanes in combination with either $\text{CB}[7]$ or $\text{CB}[8]$. For each of these host-guest species, $\text{CB}[7]$ or $\text{CB}[8]$ was demonstrated to reside in the middle of the BIPY^{2+} cations as a result of steric interactions with the anchoring platform. Upon reduction of the terminal BIPY^{2+} cations, these [3]pseudorotaxanes spontaneously

The characterization of the host-guest complexes formed between MC3^{2+} , MC4^{2+} , C23^{4+} , and C24^{4+} with $\text{CB}[8]$ was investigated by absorption spectrophotometry (Figures S29 to S32 in the ESI) and ESI-MS measurements (Figure 7). Due to the limited solubility of $\text{CB}[8]$, we were not able to perform ^1H NMR titrations of the corresponding viologens. Similarly to the [n]pseudorotaxanes formed with $\text{CB}[7]$, we were able to evidence the formation of [2]pseudorotaxanes with MC3^{2+} and MC4^{2+} (namely $\text{MC3}^{2+}\text{-(CB}[8])$ and $\text{MC4}^{2+}\text{-(CB}[8])$) and [3]pseudorotaxanes with C23^{4+} and C24^{4+} (namely a $\text{C23}^{4+}\text{-(CB}[8])_2$ and $\text{C24}^{4+}\text{-(CB}[8])_2$).

For the latter [3]pseudorotaxanes, chemical reduction of the BIPY^{2+} electroactive units led to dethreading of the $\text{CB}[8]$ macrocyclic host despite its larger cavity size (Figures S31 and S32 in the ESI). This suggests significant steric constraints with the calix[4]arene platform and/or poor flexibility of the designed systems even though longer and apparent more flexible spacers were introduced. However, the MC3^{2+} and MC4^{2+} models clearly led (Figures S29 and S30 in the ESI) to the expected [3]pseudorotaxanes ($\text{MC3}^{2+}\text{-(CB}[8])$ and ($\text{MC4}^{2+}\text{-(CB}[8])$) thus emphasizing the detrimental role of the anchoring platform rather than the length of the spacers. The occurrence of [3]pseudorotaxanes ($\text{MC3}^{2+}\text{-(CB}[8])$ and ($\text{MC4}^{2+}\text{-(CB}[8])$) in a reducing milieu is in agreement with the published reports²⁶⁻²⁸ that showed that the stability of methyl viologen dimer (MV^{2+})₂ was significantly increased in the presence of $\text{CB}[8]$ following the formation of a 2:1 (MV^{2+})₂ $\text{-(CB}[8])$ supramolecular complex in water.

dissociate as the result of a strong intramolecular dimerization of the two face-to-face viologen radical cations. The arrangement of the BIPY^{2+} radical cations within the dimeric species seemingly relies on the length of the spacer that links the electroactive units to the anchoring moiety. Dethreading and dimerization do not occur in experiments involving $\text{CB}[7]$ (or $\text{CB}[8]$) and either of the two monomeric viologen guests MC3^{2+} and MC4^{2+} , which were used as models. On the other hand, the model compounds provided unexpected and valuable properties. Thanks to the presence of an alkoxy-phenyl substitution, [3]pseudorotaxanes such as $\text{MC3}^{2+}\text{-(CB}[7])_2$ and $\text{MC4}^{2+}\text{-(CB}[7])_2$ were characterized. For the $\text{MC3}^{2+}\text{-(CB}[7])$ and

MC3²⁺⊂CB[7] [2]pseudorotaxanes, no redox-triggered translocation of **CB[7]** between the two binding stations were observed indicating that the **BIPY²⁺** unit remained the favored recognition site whatever its redox state. In the presence of **CB[8]**, similar properties were emphasized with the oxidized form of **MC3²⁺** and **MC4²⁺** threads. However, [3]pseudorotaxanes (**MC3^(••)⊂CB[8]**) and (**MC4^(••)⊂CB[8]**) were predominantly formed as a consequence of the larger cavity size of the **CB[8]** host. This preliminary study provided interesting information for the further development of functional electroactive systems and will be extended to *tetakis* analogues displaying either a cone or 1,3-alternate arrangement of the anchoring calix[4]arene unit.

- 1 a) T. Fukino, H. Yamagishi and T. Aida, *Adv. Mater.*, 2016, **29**, 1603888; b) K. Madasamy, D. Velayutham, V. Suryanarayanan, M. Kathiresan and K.C. Ho, *J. Mater. Chem. C*, 2019, **7**, 4622-4637; c) A.F. Greene, M.K. Danielson, A.O. Delawder, K.P. Liles, X. Li, A. Natraj and A. Wellen, J.C. Barnes, *Chem. Mater.*, 2017, **29**, 9498-9508; d) K.C. Ho, H.C. Lu and H.F. Yu, *RSC Smart Mater.*, 2019, **33**, 372-405; e) J. Ding, C. Zheng, L. Wang, C. Lu, B. Zhang, Y. Chen, M. Li, G. Zhai and X. Zhuang, *J. Mater. Chem. A*, 2019, **7**, 23337-23360.
- 2 K. Wadhwa, S. Nuryyeva, A.C. Fahrenbach, M. Elhabiri, C. Platas-Iglesias and A. Trabolsi, *J. Mater. Chem. C*, 2013, **1**, 2302-2307.
- 3 a) R.J. Mortimer, *Electrochim. Acta*, 1999, **44**, 2971-2981; b) L.C. Cao, M. Mou and Y. Wang, *J. Mater. Chem.*, 2009, **19**, 3412-3418; c) S. Asaftei, M. Ciobanu, A.M. Lepadatu, E. Song and U. Beginn, *J. Mater. Chem.*, 2012, **22**, 14426-14437.
- 4 M. Kuroboshi, T. Shiba, H. Tanaka, *Tetrahedron Lett.*, 2013, **54**, 3666-3668.
- 5 K. Ciepluch, N. Katir, A. El Kadib, A. Felczak, K. Zawadzka, M. Weber, B. Klajnert, K. Lisowska, A.M. Caminada, M. Bousmina, M. Bryszewska and J.P. Majoral, *Mol. Pharmaceutics.*, 2012, **9**, 448-457.
- 6 S. Berger, A. Ghicov, Y.C. Nah and P. Schmuki, *Langmuir*, 2009, **25**, 4841-4844.
- 7 D. Cummins, G. Boschloo, M. Ryan, D. Corr, S.N. Rao and D. Fitzmaurice, *J. Phys. Chem. B.*, 2000, **104**, 11449-11459.
- 8 G. De Filpo, F.P. Nicoletta and G. Chidichimo, *Chem. Mater.*, 2006, **18**, 4662-4666.
- 9 R. Cinnsealach, G. Boschloo, S. Nagaraja Rao and D. Fitzmaurice, *Sol. Energy. Mater. Sol Cells*, 1998, **55**, 215-223.
- 10 P. Bonhôte, E. Gogniat, M. Grätzel and P.V. Ashrit, *Thin Solid Films*, 1999, **350**, 269-275.
- 11 P. Sidorov, I. Desta, M. Chessé, D. Horvath, G. Marcou, A. Varnek, E. Davioud-Charvet and M. Elhabiri, *ChemMedChem*, 2016, **11**, 1339-1351.
- 12 P.M.S. Monk, *The Viologens: Physicochemical Properties, Synthesis and Applications of the Salts of 4, 4'-Bipyridine*, Chichester, 1998.
- 13 a) J. Bruinink, C.G.A. Kregting and J.J. Ponjeé, *J. Electrochem. Soc.*, 1977, **124**, 1854-1858; b) A.G. Evans, J.C. Evans and M.W. Baker, *J. Am. Chem. Soc.*, 1977, **99**, 5882-5884; c) D. Meisel, W.A. Mulac and M.S. Matheson, *J. Phys. Chem.*, 1981, **85**, 179-187; d) E. Adar, Y. Degani, Z. Goren and I. Willner, *J. Am. Chem. Soc.*, 1986, **108**, 4696-4700; e) A. Yasuda, H. Mori and J. Seto, *J. Appl. Electrochem.*, 1987, **17**, 567-573.
- 14 a) A. Trabolsi, M. Hmadeh, N. M. Khashab, D.C. Friedman, N. Humbert, M. Elhabiri, H.A. Khatib, M.E. Belowich, A. Coskun, A.M. Albrecht-Gary and J.F. Stoddart, *New J. Chem.*, 2009, **33**, 254-263. b) W. Geuder, S. Hünig and A. Suchy, *Tetrahedron*, 1986, **42**, 1665-1677.
- 15 E.M. Kosower and J.L. Cotter, *J. Am. Chem. Soc.*, 1964, **86**, 5524-5527.
- 16 E.M. Kosower and J. Hajdu, *J. Am. Chem. Soc.*, 1971, **93**, 2534-2535.
- 17 A.C. Fahrenbach, J.C. Barnes, D.A. Lanfranchi, H. Li, A. Coskun, J. J. Gassensmith, Z. Liu, D. Benítez, A. Trabolsi, W.A. Goddard, M. Elhabiri and J.F. Stoddart, *J. Am. Chem. Soc.*, 2012, **134**, 3061-3072.
- 18 Y. Wang, J. Sun, Z. Liu, M. S. Nassar, Y.Y. Botros and J.F. Stoddart, *Chem. Sci.*, 2017, **8**, 2562-2568.
- 19 a) A.F. Greene, M.K. Danielson, A.O. Delawder, K.P. Liles, X. Li, A. Natraj, A. Wellen and J.C. Barnes, *Chem. Mater.*, 2017, **29**, 9498-9508. b) Z. Qian, X. Huang, Q. Wang, *Dyes & Pigment.*, 2017, **145**, 365-370.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 20 G.C. Granqvist, *Handbook of Inorganic Electrochromic Materials*, Elsevier, Amsterdam, 1995.
- 21 P.M.S. Monk, R.J. Mortimer and D.R. Rosseinsky, *Electrochromism and Electrochromic Devices*, Cambridge University Press, Cambridge, UK, 2007.
- 22 J.R.J.J. Platt, *J. Chem. Phys.*, 1961, **34**, 862-863.
- 23 J.W. Park, N.H. Choi and J.H. Kim, *J. Phys. Chem.*, 1996, **100**, 769-774.
- 24 P.A. Quintela and A.E. Kaifer, *Langmuir*, 1987, **3**, 769-773.
- 25 A. Trabolsi, N. Khashab, A.C. Fahrenbach, D.C. Friedman, M.T. Colvin, K.K. Cotí, D. Benítez, E. Tkatchouk, J.C. Olsen, M.E. Belowich, R. Carmielli, H.A. Khatib, W.A. Goddard III, M.R. Wasielewski and J.F. Stoddart, *Nature Chem.*, 2010, **2**, 42-49.
- 26 H.J. Kim, W.S. Jeon, Y.H. Ko and K. Kim, *Proc. Natl. Acad. Sci. USA*, 2002, **99**, 5007-5011.
- 27 W. Ong, M. Gomez-Kaifer and A.E. Kaifer, *Org. Lett.*, 2002, **4**, 1791-1794.
- 28 W.S. Jeon, H.J. Kim, C. Lee and K. Kim, *Chem. Commun.*, 2002, 1828-1829.
- 29 J.W. Lee, S. Samal, N. Selvapalam, H.J. Kim and K. Kim, *Acc. Chem. Res.*, 2003, **36**, 621-630.
- 30 E. Pazos, P. Novo, C. Peinador, A. E. Kaifer and M. D. García, *Angew. Chem. Int. Ed.* 2019, **58**, 403-416.
- 31 a) K. Nchimi Nono, P. Dalvand, K. Wadhwa, S. Nuryyeva, S. Alneyadi, A. Fahrenbach, J.C. Olsen, Z. Asfari, C. Platas-Iglesias, M. Elhabiri and A. Trabolsi, *Chem. Eur. J.*, 2014, **20**, 7334-7344; b) K. Wadhwa, S. Nuryyeva, A.C. Fahrenbach, M. Elhabiri, C. Platas-Iglesias, A. Trabolsi, *J. Mater. Chem. C*, 2013, **1**, 2302-2307.
- 32 R. Kannappan, C. Bucher, E. Saint-Aman, J.C. Moutet, A. Milet, M. Oltean, E. Métay, S. Pellet-Rostaing, M. Lemaire and C. Chaix, *New J. Chem.*, 2010, **34**, 1373-1386.
- 33 W. Geuder, S. Huenig and A. Suchy, *Tetrahedron*, 1986, **42**, 1665-1667.
- 34 S.J. Atherton, K. Tsukahara and R.G. Wilkins, *J. Am. Chem. Soc.*, 1986, **108**, 3380-3385.
- 35 W.S. Abdul-Hassan, D. Roux, C. Bucher, S. Cobo, F. Molton, E. Saint-Aman and G. Royal, *Chem. Eur. J.*, 2018, **24**, 12961-12969.
- 36 M. Berville, S. Choua, C. Gourlaouen, C. Boudon, L. Ruhlmann, C. Bailly, S. Cobo, E. Saint-Aman, J.A. Wytko and J. Weiss, *ChemPhysChem*, 2017, **18**, 796-803.
- 37 M. Berville, L. Karmazin, J.A. Wytko and J. Weiss, *Chem. Commun.*, 2015, **51**, 15772-15775.
- 38 A. Iordache, M. Retegan, F. Thomas, G. Royal, E. Saint-Aman and C. Bucher, *Chem. Eur. J.*, 2012, **18**, 7648-7653.
- 39 S. Chowdhury, Y. Nassar, L. Guy, D. Frath, F. Chevallier, E. Dumont, A.P. Ramos, G.J.F. Demets and C. Bucher, *Electrochim. Acta*, 2019, **316**, 79-92.
- 40 A. Iordache, M. Oltean, A. Milet, F. Thomas, E. Saint-Aman and C. Bucher, *J. Am. Chem. Soc.*, 2012, **134**, 2653-2671.
- 41 M. Mohammad, *Electrochim. Acta*, 1988, **33**, 417-419.
- 42 T.G. Zhan, T.Y. Zhou, F. Lin, L. Zhang, C. Zhou, Q.Y. Qi, Z.T. Li and X. Zhao, *Org. Chem. Front.*, 2016, **3**, 1635-1645.
- 43 K. Madasamy and M. Kathiresan, *ChemistrySelect*, 2016, **1**, 354-359.
- 44 Y. Wang, M. Frasconi, W.G. Liu, Z. Liu, A.A. Sarjeant, M.S. Nassar, Y.Y. Botros, W.A. Goddard and J.F. Stoddart, *J. Am. Chem. Soc.*, 2015, **137**, 876-885.

- 45 J. Iehl, M. Frasconi, H.P. Jacquot de Rouville, N. Renaud, S.M. Dyar, N.L. Strutt, R. Carmieli, M.R. Wasielewski, M.A. Ratner, J.F. Nierengarten and J.F. Stoddart *Chem. Sci.*, 2013, **4**, 1462-1469.
- 46 a) A. Iordache, R. Kanappan, E. Métay, M.C. Duclos, S. Pellet-Rostaing, M. Lemaire, A. Milet, E. Saint-Aman and C. Bucher, *Org. Biomol. Chem.*, 2013, **11**, 4383-4389; b) C. Kahlfuss, E. Métay, M. C. Duclos, M. Lemaire, M. Oltean, A. Milet, E. Saint-Aman and C. Bucher, *C. R. Chimie*, 2014, **17**, 505-511.
- 47 C. Kahlfuss, E. Metay, M.C. Duclos, M. Lemaire, A. Milet, E. Saint-Aman and C. Bucher, *Chem. Eur. J.*, 2015, **21**, 2090-2106.
- 48 M. Elanchezian, K. Theyagarajan, D. Saravanakumar, K. Thenmozhi and S. Senthilkumar, *Mater. Today Chem.*, 2020, **16**, 100274.
- 49 C. M. Ronconi, J. F. Stoddart, V. Balzani, M. Baroncini, P. Ceroni, C. Giansante and M. Venturi, *Chem. Eur. J.* 2008, **14**, 8365-8373.
- 50 G. Das, T. Prakasam, S. Nuryyeva, D. S. Han, A. Abdel-Wahab, J.C. Olsen, K. Polychronopoulou, C. Platas-Iglesias, F. Ravaux, M. Jouiad and A. Trabolsi, *J. Mater. Chem. A*, 2016, **4**, 15361-15369.
- 51 M. Marchini, M. Baroncini, G. Bergamini, P. Ceroni, M. D'Angelantonio, P. Franchi, M. Lucarini, F. Negri, T. Szreder, M. Venturi, *Chem. Eur. J.* 2017, **23**, 6380-6390
- 52 I. Thondorf, A. Shivanyuk and V. Bohmer, In *Calixarene 2001*, Z. Asfari, V. Bohmer, J. Harrowfield and J. Vicens, Eds., Kluwer Academic publisher, Dordrecht, 2001, **2**.
- 53 *Functional Synthetic Receptors*, T. Schrader and A.D. Hamilton, Eds., Wiley-VCH, Weinheim, 2005.
- 54 K. Ariga and T. Kunitake *Supramolecular Chemistry-Fundamentals and Applications*, Springer, Berlin, 2005.
- 55 *Highlights in Bioorganic Chemistry: Methods and Applications*, C. Schmuck and H. Wennemers, Eds., Wiley-VCH, Weinheim, 2004.
- 56 S. Kubik, C. Reyheller and S. Stuwe, *J. Incl. Phenom. Macro. Chem.*, 2005, **52**, 137-187.
- 57 H.C. Ko, S.A. Park, W.K. Paik, H. Lee, *Synth. Met.*, 2002, **132**, 15-20.
- 58 a) S. Luan, Q.Ge, Y. Chen, M. Dai, J. Yang, K. Li, D. Liu, L. Zhao, *Bioorg. Med. Chem. Lett.*, 2017, **27**, 1943-1948. b) W. Wang, C. Sheng, X. Che, H. Ji, Y. Cao, Z. Miao, J. Yao, W. Zhang, *Bioorg. Med. Chem. Lett.*, 2009, **19**, 5965-5969.
- 59 W.-W. Gu, W.-J. Chen, C.-G. Yan, *Supramol. Chem.*, 2015, **27**, 407-413.
- 60 R. Pomecko, PhD Thesis University of Strasbourg, 2007.
- 61 Z.T. Li, G.Z. Ji, C.X. Zhao, S.D. Yuan, H. Ding, C. Huang, A.L. Du and M. Wei, *J. Org. Chem.*, 1999, **64**, 3572-3584.
- 62 Y. Zhao and D. G. Truhlar, *Theor. Chem. Acc.*, 2008, **120**, 215-241.
- 63 Gaussian 16, Revision B.01, M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman, and D. J. Fox, Gaussian, Inc., Wallingford CT, 2016.
- 64 a) G. Scalmani and M. J. Frisch, *J. Chem. Phys.*, 2010, **132**, 114110; b) J. Tomasi, B. Mennucci and R. Cammi, *Chem. Rev.*, 2005, **105**, 2999-3094.
- 65 A. S. Jalilov, S. Patwardhan, A. Singh, T. Simeon, A. A. Sarjeant, G. C. Schatz and F. D. Lewis, *J. Phys. Chem. B*, 2014, **118**, 125-133.
- 66 A. Alberti, In *Organic Photochromic and Thermochromic Compounds*, J.C. Crano and R.J. Guglielmetti, Eds., Kluwer Academic publisher, 2002, **4**, 211-240.
- 67 W. L. Mock and N.-Y. Shih, *J. Org. Chem.*, 1986, **51**, 4440-4446.
- 68 B. Perlmutter-Hayman, *Acc. Chem. Res.*, 1986, **19**, 90-96.
- 69 C. Lee, M.S. Moon and J.W. Park, *J. Electroanal. Chem.*, 1996, **407**, 161-167.
- 70 J.F. Stargardt and F.M. Hawkrige, *Anal. Chim. Acta*, 1983, **146**, 1-8.
- 71 P.M.S. Monk, N.M. Hodgkinson and S.A. Ramzan, *Dyes Pigm.*, 1999, **43**, 207-217.