



New fluorescent tetraphenylporphyrin-based dendrimers with alkene-linked fluorenyl antennae designed for oxygen sensitization

D. Yao, L. Shi, Z. Sun, Mireille Blanchard-desce, O. Mongin, Frédéric Paul, C.O. Paul-Roth

► To cite this version:

D. Yao, L. Shi, Z. Sun, Mireille Blanchard-desce, O. Mongin, et al.. New fluorescent tetraphenylporphyrin-based dendrimers with alkene-linked fluorenyl antennae designed for oxygen sensitization. *Comptes Rendus. Chimie*, 2021, 24 (S3), pp.57-70. 10.5802/crchim.99 . hal-03334340

HAL Id: hal-03334340

<https://hal.science/hal-03334340>

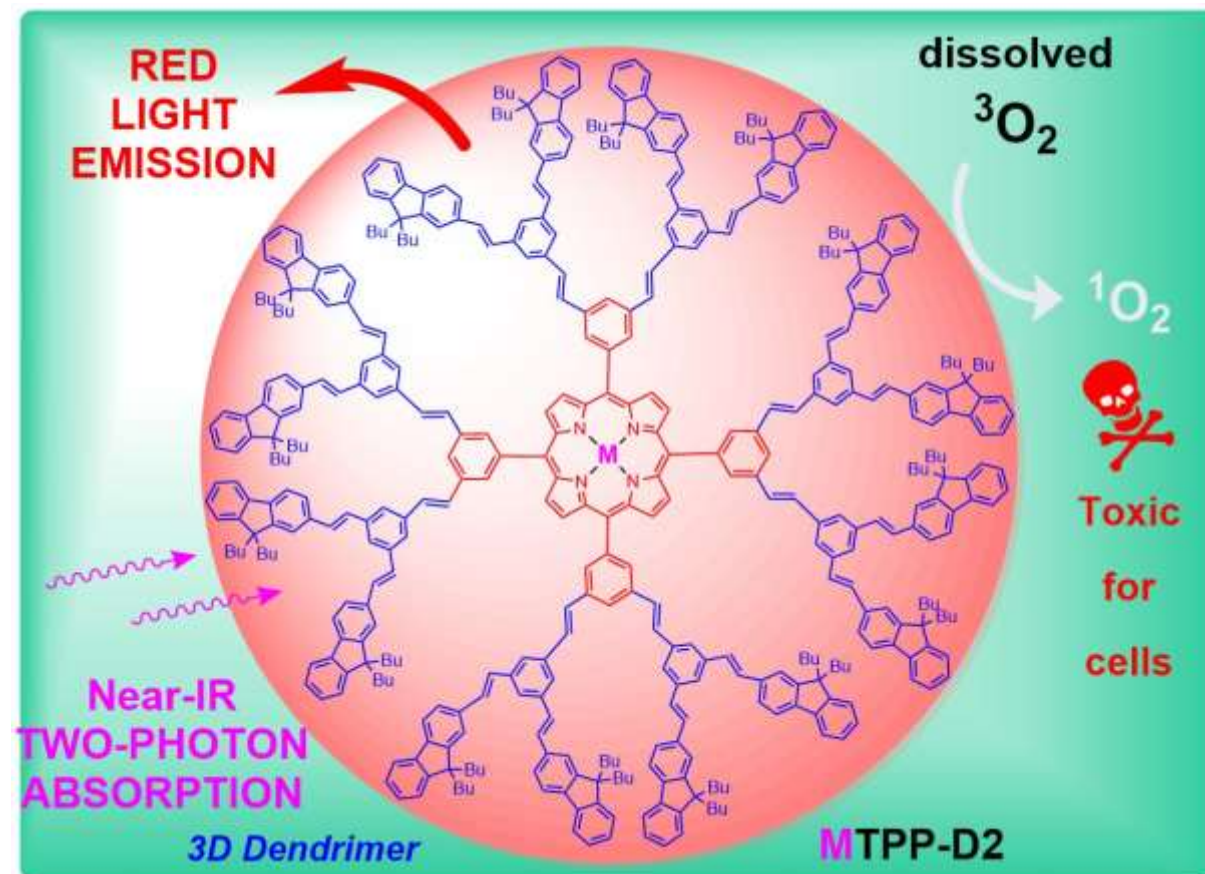
Submitted on 19 Oct 2021

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

New Fluorescent Tetraphenylporphyrin-Based Dendrimers with Alkene-Linked Fluorenyl Antennae designed for Oxygen Sensitization

Dandan Yao, Limiao Shi, Zhipeng Sun, Mireille Blanchard-Desce, Olivier Mongin, Frédéric Paul, Christine O. Paul-Roth



A new family of conjugated porphyrin-based dendrimers with fluorenyl arms has been synthesized. Their detailed luminescence properties and selected photophysical properties are discussed in the context of two-photon-induced theranostics.

New Fluorescent tetraphenylporphyrin-based Dendrimers with Alkene-Linked Fluorenyl Antennae designed for Oxygen Sensitization

Dandan Yao,^a Limiao Shi,^a Zhipeng Sun,^a Mireille Blanchard-Desce,^b
Olivier Mongin,^a Frédéric Paul,^a Christine O. Paul-Roth,^{*,a}

^a Univ Rennes, INSA Rennes, CNRS, ISCR (Institut des Sciences Chimiques de Rennes) – UMR 6226, F-35000 Rennes, France

^b Université de Bordeaux, Institut des Sciences Moléculaires (CNRS UMR 5255), 33405 Talence, France

*Corresponding author: christine.paul@univ-rennes1.fr or christine.paul@insa-rennes.fr

Tel: (+33) (0) 2 23 23 63 72; Fax: (+33) (0) 2 23 23 63 72

Abstract. The design of porphyrin-based dendrimers featuring conjugated fluorenyl dendrons *via* alkene spacers allows evaluating the importance of the role of these spacers on the optical properties of interest. In the continuation of previous studies, a second-generation porphyrin-based dendrimer was synthesized and metalated by Zn(II) along with its known first-generation homologue. The targeted free-base porphyrin was obtained by repetitively cycling a Michaelis-Arbuzov reaction and a Horner-Wadsworth-Emmons reaction to construct the desired vinyl-containing dendrons. After metalation by Zn(II), *meso*-tetraphenylporphyrin-cored dendrimers with eight (**ZnTPP-D1**) and sixteen (**ZnTPP-D2**) fluorenyl arms at their periphery were eventually isolated. These species allow evaluating the influence of the central metal and of the 1,2-alkyne for 1,2-alkene exchange on fluorescence, oxygen photosensitization and two-photon absorption. Such structure-property relationships are currently needed for the design of optimal dendrimeric photosensitizers allowing combined two-photon-based photodynamic therapy (2P PDT) and imaging.

Keywords. Porphyrin, fluorenyl, fluorescence, energy transfer, double bond, dendrimer, photodynamic therapy

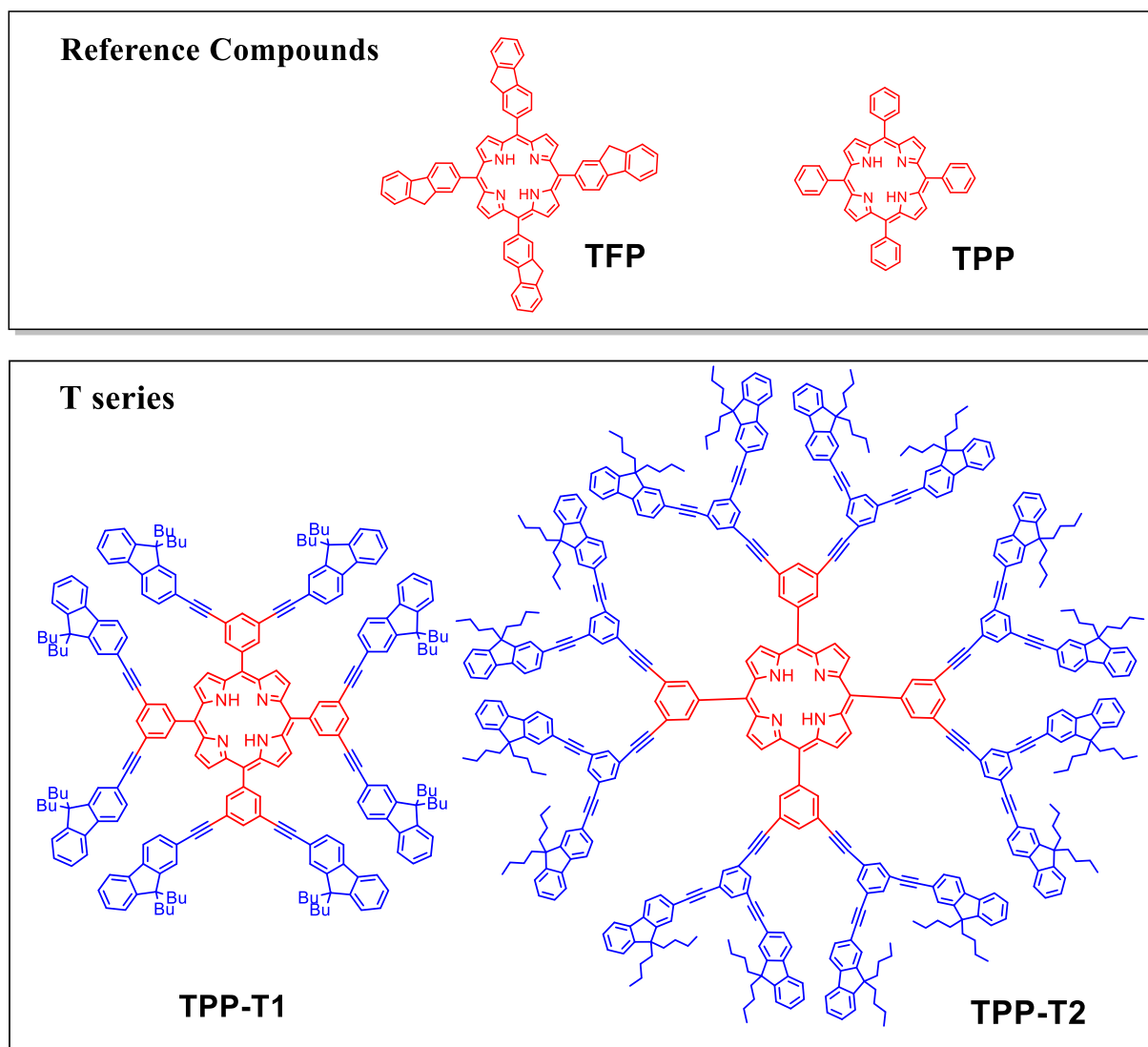
1. Introduction

There has been great interest in porphyrin systems because the peripheral substituents on the macrocyclic core allows to significantly modulate the photophysical properties at will. Thus, light-harvesting compounds can be easily obtained by attaching four strongly absorbing energy-donor dendrons at the *meso* positions of the central porphyrin core which will play the role of peripheral “antenna” for the porphyrin.¹ In this respect, Fréchet and coworkers² have reported porphyrin systems with fluorene-containing oligoether-type dendrons as efficient one- and two-photon light-harvesting units and demonstrated that such an “antenna effect” was facilitated in dendritic architectures *versus* linear architectures.^{2b-c} More recently, related star-shaped porphyrins with fully conjugated oligofluorene arms have also been reported by Bo and coworkers and shown to behave as a remarkable light-harvesting systems.³

In this context, we have previously reported the synthesis of porphyrin possessing four fluorenyl arms directly connected at the *meso*-positions. This compound (**TFP**; Figure 1a).⁴ exhibited a high luminescence quantum yields (24%), demonstrating the good capacity of the fluorenyl units to enhance emission by increasing the radiative process.⁵ Subsequently, we synthesized two series of substituted *meso*-porphyrin dendrimers with terminal fluorenyl arms, taking **TFP** as the central unit: (i) a non-conjugated family with flexible ether linkage⁶ and, more recently, (ii) a conjugated family with rigid alkynyl linkage (Figure 1b).⁷ We then could show that these molecular architectures were promising models for the design of new biphotonic photosensitizers for theranostics *i.e.* allowing to perform photodynamic therapy⁸ and fluorescence imaging after two-photon excitation.⁷ Due to the practical advantages of two-photon excitation, this field has attracted attention and is rapidly expanding, several porphyrin-based such photosensitizers having been reported to date.^{9,10,11} In order to gain additional insight about the potential of TFP-cored dendrimers such as **TFP-T1** or **TFP-T2** in this field, we have started systematically investigating the impact of various structural changes taking place in the peripheral light-harvesting antenna on the photophysical properties of these systems.

Given that 1,2-alkene spacers are known to promote electronic communication better than 1,2-alkyne ones,¹² we recently started exploring the optical properties of related dendrimers featuring alkene linkages. However this was exclusively done for the *first generation* dendrimer **TFP-D1** (Figure 1c; M = 2H) resulting in a significant improvement in the linear and nonlinear optical properties *vs.* **TFP-T1**.¹³ This statement prompted us to test higher generation

dendrimers of this kind. Accordingly, we now report our efforts to isolate the *second generation* dendrimer **TPP-D2** and the metalated species **ZnTPP-D1** and **ZnTPP-D2**. Subsequently, their linear and nonlinear optical properties of interest will be disclosed as well as evidence for the very efficient energy-transfer process taking place from the conjugated dendrons to the porphyrin core in these new species. These data will then be compared to those of their known alkyne-containing analogues (or T series: Figure 1b).



D series (M = 2H, Zn)

MTPP-D1

MTPP-D2

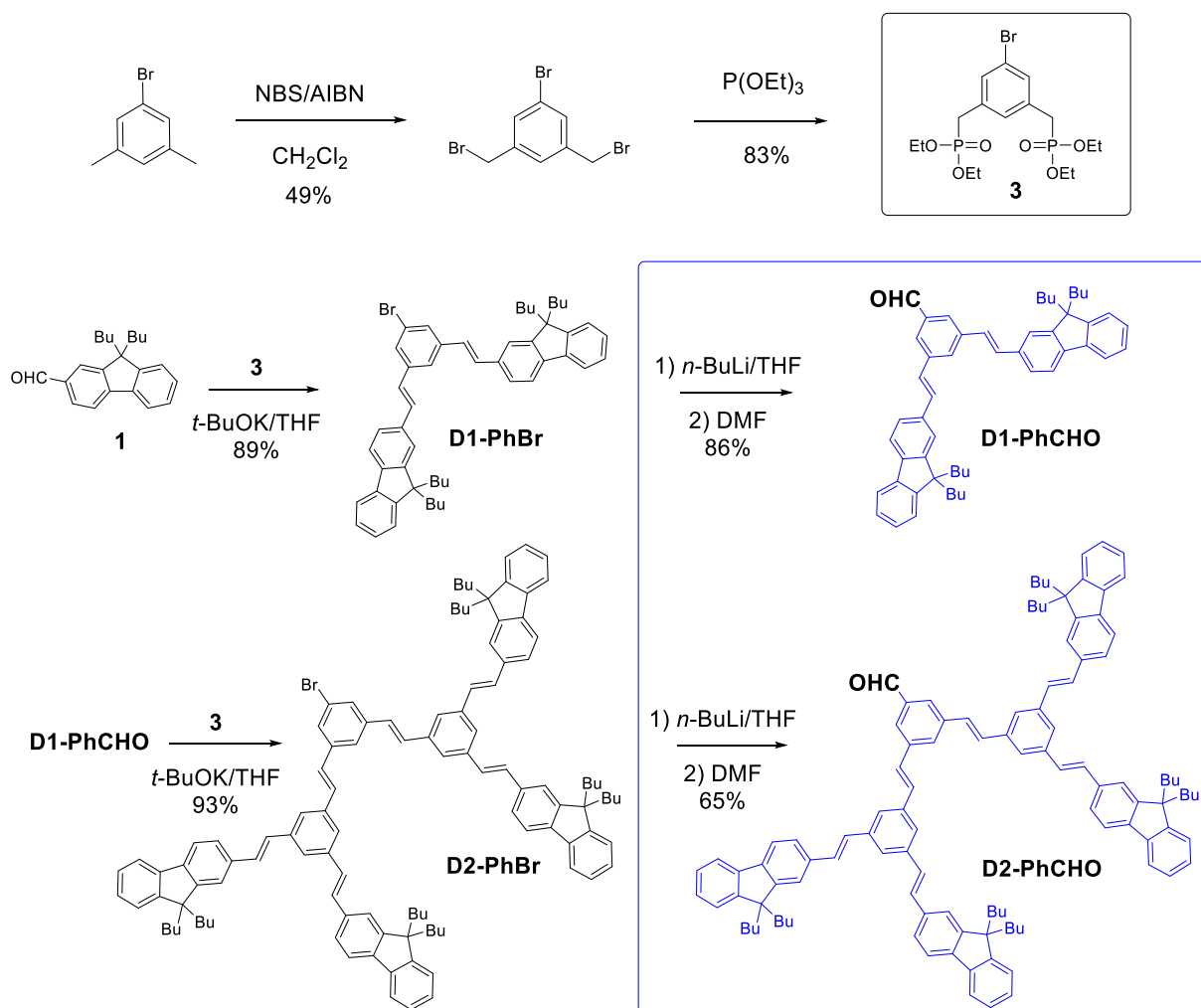
Figure 1 a) Reference compounds; b) Previously reported alkynyl-bridged free-base porphyrin dendrimers (**T series**) based on **TPP** (**TPP-T1** and **TPP-T2**); c) New porphyrin dendrimers (**D series**) based on **TPP** (**MTPP-D1** and **MTPP-D2**; M = 2H, Zn).

2. Results and Discussion

2.1. Synthesis and characterization

This new family of dendrimers was prepared the generation of the required dendrons (**D1-PhCHO** and **D2-PhCHO**), followed by their condensation with pyrrole to give the desired free-base porphyrins as intermediates which were then metalated by Zn(II) to give the final dendrimers **ZnTPP-D1** and **ZnTPP-D2**.

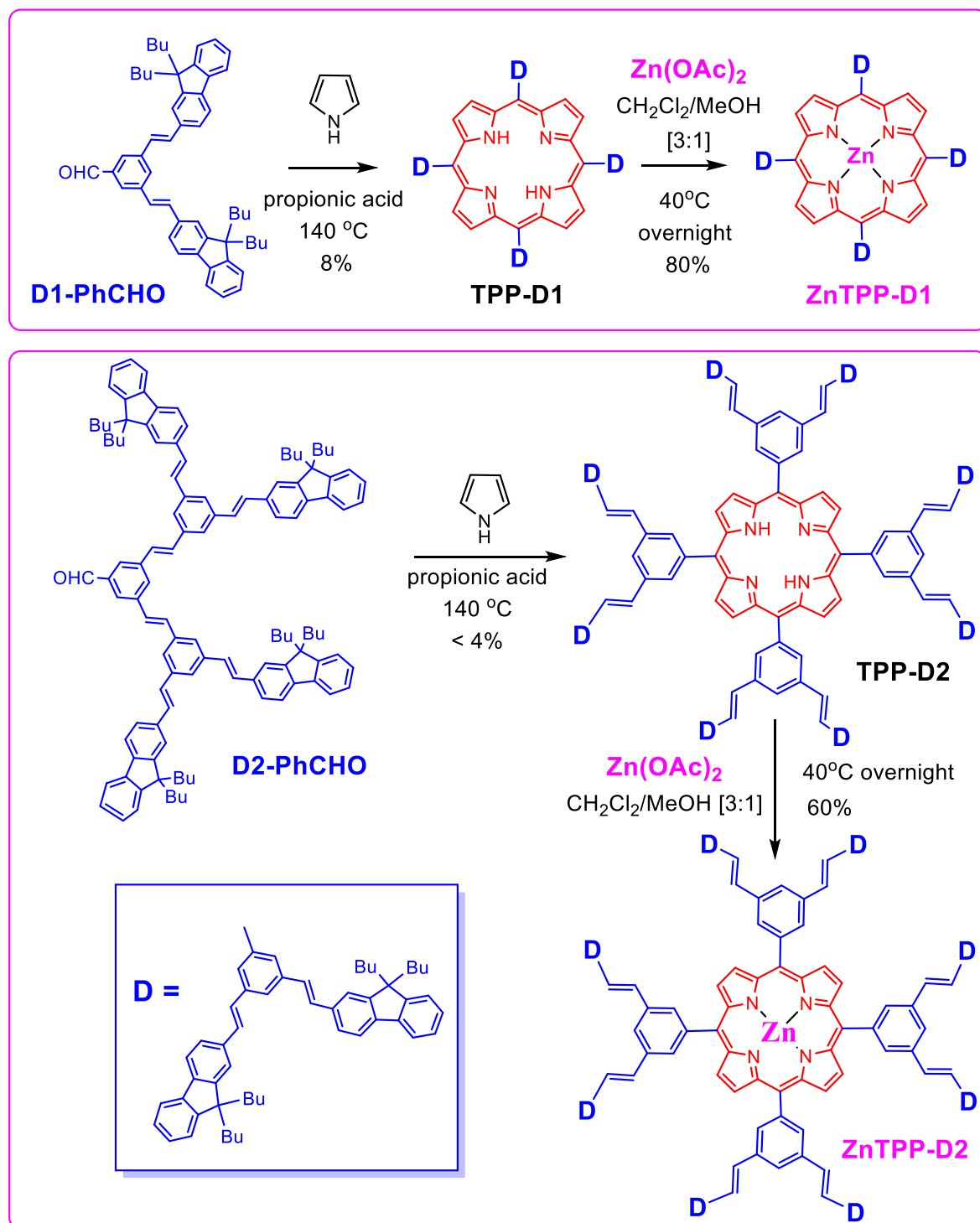
Dendron synthesis. The synthesis of the two generations of vinyl-bridged aldehyde-terminated dendrons **D1-PhCHO** and **D2-PhCH** is described in Scheme 1. First, commercial 1-bromo-3,5-bismethylbenzene was halogenated with benzyl by *N*-bromosuccinimide (NBS) using azobis-isobutyronitrile (AIBN) as radical initiator. Given that this bromination takes usually place non selectively,¹⁴⁻¹⁵ the reaction conditions were optimized (temperature, time and solvent). The resulting conditions gave mainly the desired dibrominated product along with mono- and tri-brominated byproducts. The former byproduct could be isolated by chromatography (heptane), while the latter could not be fully separated from targeted dibrominated product (ratio tri/di of 1/4 by ¹H NMR). In the next step, this mixture containing 49% of desired product was reacted directly with excess of P(OEt)₃ under reflux following a Michaelis-Arbuzov¹⁴ protocol. The desired product **3** and its bromo-substituted byproduct were both formed and were subsequently separated by chromatography (Scheme 1). The unwanted byproduct could be easily eluted using CH₂Cl₂, while the target product **3** was collected using ethyl acetate as a colorless oil in 83% yield. After a subsequent Horner-Wadsworth-Emmons (HWE) reaction,¹⁴ the compound **3** reacts with the previously prepared aldehyde **1**^{7a} in presence of *t*-BuOK/THF to give the double-bonded precursor **D1-PhBr** in 89% yield. The aldehyde **D1-PhCHO** can then be obtained in two steps from this product in 86% yield. Repetition of the HWE reaction¹⁴ between this new aldehyde and **3** gave access to the *second generation* intermediate **D2-PhBr** in 93% yield, and subsequently to the corresponding **D2-PhCHO** dendron in 65% yield (Scheme 1).



Scheme 1. Synthetic routes for alkene-bridged dendrons **D1-PhCHO**¹³ and **D2-PhCHO**.

Porphyrin synthesis. Two synthetic methods are most often used to synthesize porphyrins: the Adler-Longo¹⁶ or the Lindsey reaction.¹⁷ Both are efficient for synthesizing porphyrins substituted at their *meso* positions. Given that the *first generation* dendrimer **TPP-D1** was previously isolated using the Adler-Longo approach,¹³ these reaction conditions were used again to synthesize **TPP-D2**. This compound could however not be isolated pure after several purification attempts by chromatography and subsequent recrystallizations (CHCl_3 and MeOH), the yield of crude **TPP-D2** in the isolated solid being below 4%. Fortunately all these porphyrin dendrimers have good solubilities in common organic solvents, allowing for their easy metalation and this approach provided a way to access the pure zinc complex from the mixture in the case of **ZnTPP-D2** (Scheme 2). Thus, the corresponding zinc complex **ZnTPP-D1** was formed at 40 °C overnight from **TPP-D2** using Zn(OAc)_2 in $\text{CH}_2\text{Cl}_2/\text{MeOH}$ and isolated

pure in 60% yield (Scheme 2). Using similar conditions for metalation **ZnTPP-D1** was isolated in 80% yield from **TPP-D1**.

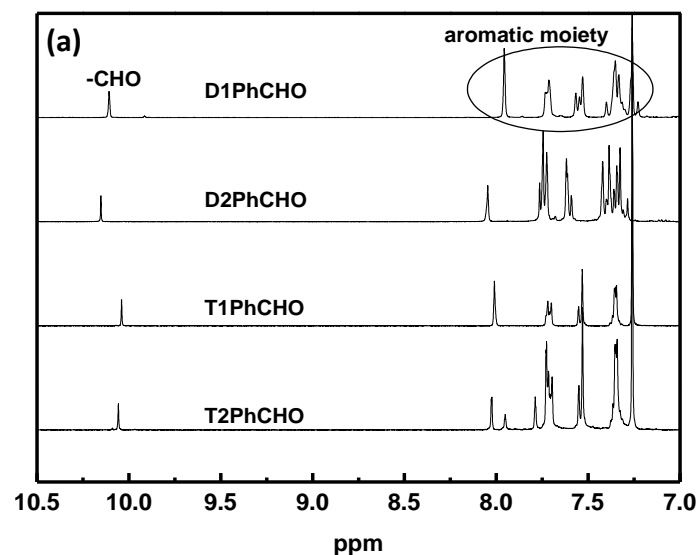


Scheme 2. Synthesis of vinyl-bridged porphyrin dendrimers (**D** series) based on **TPP**-cored porphyrin (**TPP-D1** and **TPP-D2**) and corresponding zinc(II) complexes (**ZnTPP-D1** and **ZnTPP-D2**).

From a purely synthetic standpoint, the isolated yields in alkene-bridged dendrimers (**D**-

type series; Figure 1c) were always lower than those of their alkyne-bridged analogues (T-type series; Figure 1b), although rigorously similar reaction conditions have been used (the isolated yields of **TPP-T1** and **TPP-T2** were 18% and 13%, respectively.⁷), making the D-type series dendrimers more challenging to obtain via the Adler-Longo approach.^{7,18}

¹H NMR analysis. The aldehyde dendrons **D1-PhCHO** and **D2-PhCHO**, and the corresponding dendrimers, *i.e.* the free-base and metalated porphyrins **MTPP-D1** and **MTPP-D2**, were characterized by ¹H NMR analysis (Figures 2-4). The Figure 2 shows the full ¹H NMR spectra of the dendrons compared to those of their analogues with triple bonds (**T1-PhCHO** and **T2-PhCHO**). They all show three diagnostic signatures: (i) the aldehyde proton as a singlet, around 10 ppm; (ii) the aromatic protons located as multiplets, in region 7-8 ppm, belonging to protons of phenyl and fluorenyl, partially identified; (iii) four groups of alkyl protons H_{a,b,c,d} located at 0.5-2.1 ppm are assigned to *n*-butyl of fluorenyl. We can particularly notice that for double bonded dendrons **D1-PhCHO** and **D2-PhCHO**, we observe four additional alkene protons as a broad peak around 7.18-7.40 ppm.



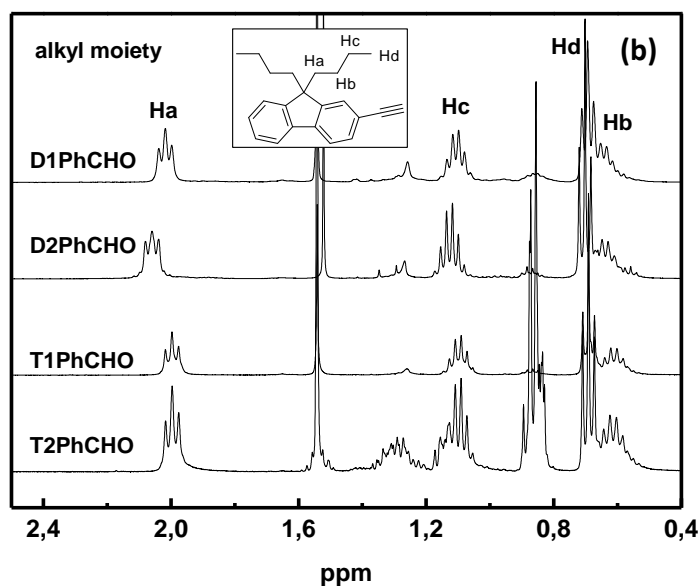


Figure 2. Aromatic (a) and Alkyl (b) moieties ^1H NMR spectra (400 MHz) of **D-style** dendrons **D1-PhCHO** and **D2-PhCHO** compared to reported **T-style** dendrons **T1-PhCHO** and **T2-PhCHO** in CDCl_3 .⁷

The full ^1H NMR spectra in CDCl_3 of the corresponding free-bases porphyrin **TPP-D1** (Figure 3) show four diagnostic signatures: (i) the β -pyrrolic protons of the porphyrin core (H_β) around 9 ppm, (ii) the aromatic protons around 7.3-8.4 ppm, (iii) the alkyl protons of the various butyl chains around 2.2-0.5 ppm and (iv) the NH protons of porphyrin cavity around -2.6 ppm. For **TPP-D1** as for reported **TPP-T1** and **TPP-T2**,⁷ we observe eight protons H_β of porphyrin ring located around 9 ppm. For aromatic and vinylic protons (around 7.2-8.4 ppm), only some of them can be easily assigned like H_A and H_B . Again, the vinyl protons of **TPP-D1** give rise to a broad peak around 7.2-7.3 ppm. For all these dendrimers, the *n*-butyl protons ($\text{H}_{\text{a,b,c,d}}$) are simply assigned to four groups of signals located at 0.5-2.1 ppm as for the corresponding dendrons. In contrast, for the larger dendrimer **TPP-D2**, a ^1H NMR spectrum with broad signals was obtained (see ESI) and only for the corresponding zinc(II) complex **ZnTPP-D2** was the spectrum well resolved (Figure 4). Some characteristic signals are also readily identified; (i) eight protons H_β of porphyrin ring located around 9 ppm, (ii) H_A and H_B in the phenyl arms. However, most of the aromatic and vinylic protons overlap (7.2-8.4 ppm) because of the larger molecular structure.

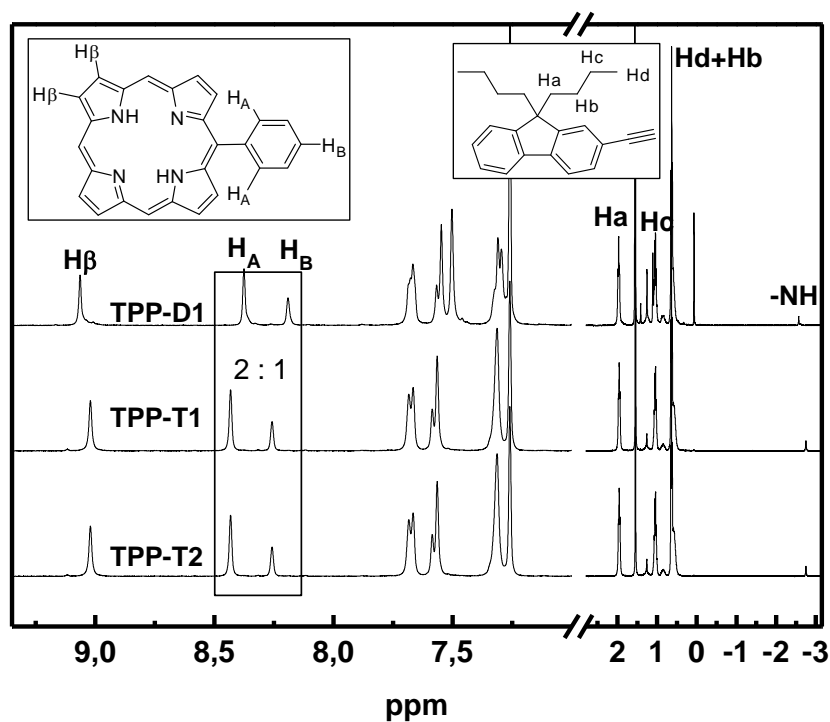


Figure 3. Full ^1H NMR spectra (400 MHz) of the free-base TPP-cored dendrimers **TPP-D1**, **TPP-T1** and **TPP-T2** in CDCl_3 .

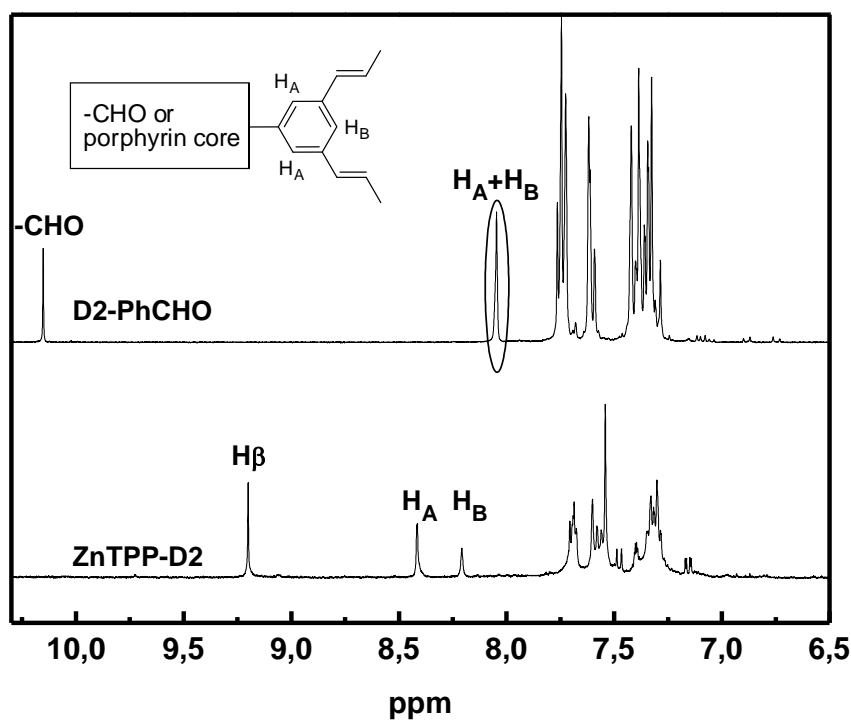


Figure 4. Partial ^1H NMR spectra (400 MHz) of the dendron **D2-PhCHO** and of the corresponding zinc(II) complex **ZnTPP-D2** in CDCl_3 .

2.2. Photophysical properties

UV-visible absorption and emission spectra, after excitation in the Soret-band,¹⁹ were recorded for the isolated dendrimers at room temperature (Table 1 and Figures 5 and 6). The free-base tetraphenylporphyrin (**TPP**; Figure 1a) and the corresponding zinc complex (**ZnTPP**) were chosen as references compounds. Their two-photon oxygen photosensitizing yields were subsequently evaluated and compared to those of their alkyne analogues (**TPP-T1** and **TPP-T2**) to analyze the impact of this structural modification (Tables 2-3).

For the free-base porphyrins series, all absorption spectra are typical of porphyrin derivatives with an intense Soret-band around 430 nm and four Q-bands in the 520-650 nm range.^{19a} Compared to **TPP**, the dendrimers **TPP-D1**, **TPP-T1** and **TPP-T2** present an additional absorption around 300-400 nm, corresponding to the conjugated fluorenyl dendrons. This band is of almost similar intensity for **TPP-D1** and **TPP-T1**, most likely in reason of their identical terminal fluorenyl number, however, a red-shift is observed for **TPP-D1** compared to **TPP-T1**, likely due to the improved conjugation between core and arms in the former compound. Concerning the Soret band, the vinyl-bridged dendrimer **TPP-D1** present also a larger red-shift (15 nm) than its alkyne analogue **TPP-T1** (9 nm) *versus* the Soret band of the reference **TPP**, which reflect the extension of the porphyrin π manifold²⁰ through conjugation with the peripheral arms at the *meso* positions.

The corresponding zinc complexes **ZnTPP-D1** and **ZnTPP-D2** exhibit characteristic changes in their electronic spectra compared to those of the corresponding the free-base porphyrins **TPP-D1** and **TPP-D2**.^{19a,21} Only two Q-band absorption are now observed around 552 nm and 635 nm due to the symmetry change from D_{2h} to D_{4h} symmetry upon metalation and an intense Soret band around 435 nm (Table 1 and Figure 6). An additional broad band is also observed in UV range (346 nm) which corresponds to π - π^* absorption of the fluorenyl chromophores. This UV absorption is weaker for the *first generation* zinc(II) complex (**ZnTPP-D1**) than for the higher generation dendrimer (**ZnTPP-D2**) due to the smaller number (eight vs. sixteen) of fluorenyl groups present in the peripheral arms.

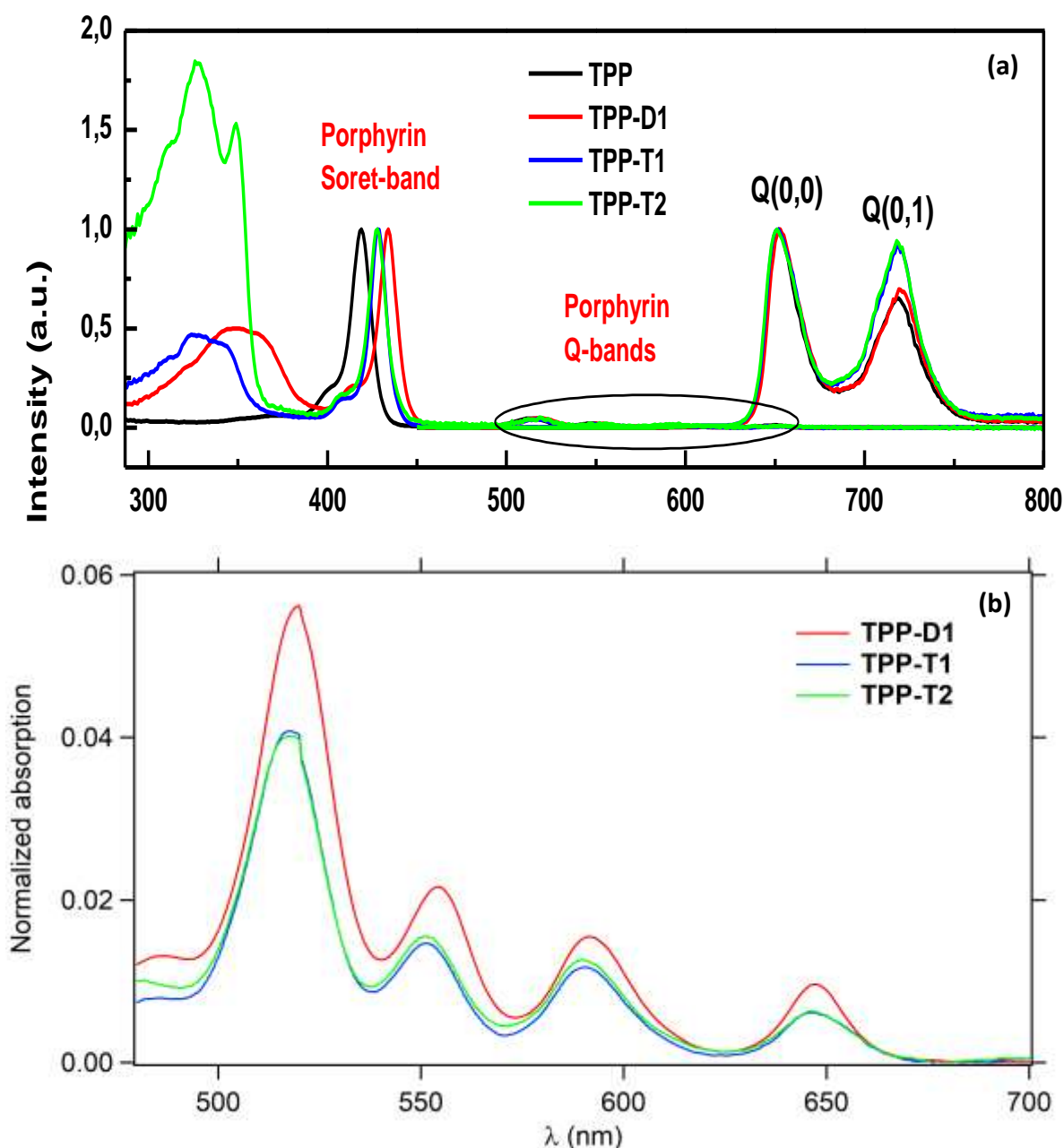


Figure 5. Normalized absorption and emission spectra of selected TPP-cored dendrimers (**TPP-D1**, **TPP-T1** and **TPP-T2**) compared to reference **TPP** in toluene (a). Detail of the Q bands of **TPP-D1**, **TPP-T1** and **TPP-T2** (b).

Upon excitation in their Soret-band, all these compounds exhibit the characteristic porphyrin emission peaks Q(0,0) and Q(0,1).^{5,20} After normalizing their emission spectra on their Q(0,0) peaks, all these compounds exhibit two emission peaks at similar wavelengths, but with different intensities (Figure 5). Compared to **TPP**, the Q(0,1) band of all dendrimers does not change in intensity for **TPP-D1** but increases for **TPP-T1** and **TPP-T2**. The emission spectra of zinc(II) porphyrin complexes usually consist of three sub-bands assigned to a vibronic progression from a Q state: Q(0,0), Q(0,1) and weak Q(0,2), the last one, near 720 nm

being usually too weak to be observed.²² Presently, with **ZnTPP-D1** and **ZnTPP-D2** exhibits the two expected Q-bands around 603 nm and 653 nm (Figure 6), the blue shift of these bands compared to the corresponding free-bases being ascribable to the metal coordination. The quantum yields (Φ_F) were also measured for all these compounds (Table 1). While the free-base porphyrin-based dendrimers have similar quantum yields ($\Phi_F = 12$ -13%) than the **TPP** ($\Phi_F = 11\%$), these values drop drastically (3-6%) when a metal like zinc(II) is introduced in the porphyrin cavity, as also observed for the reference **TPP**.

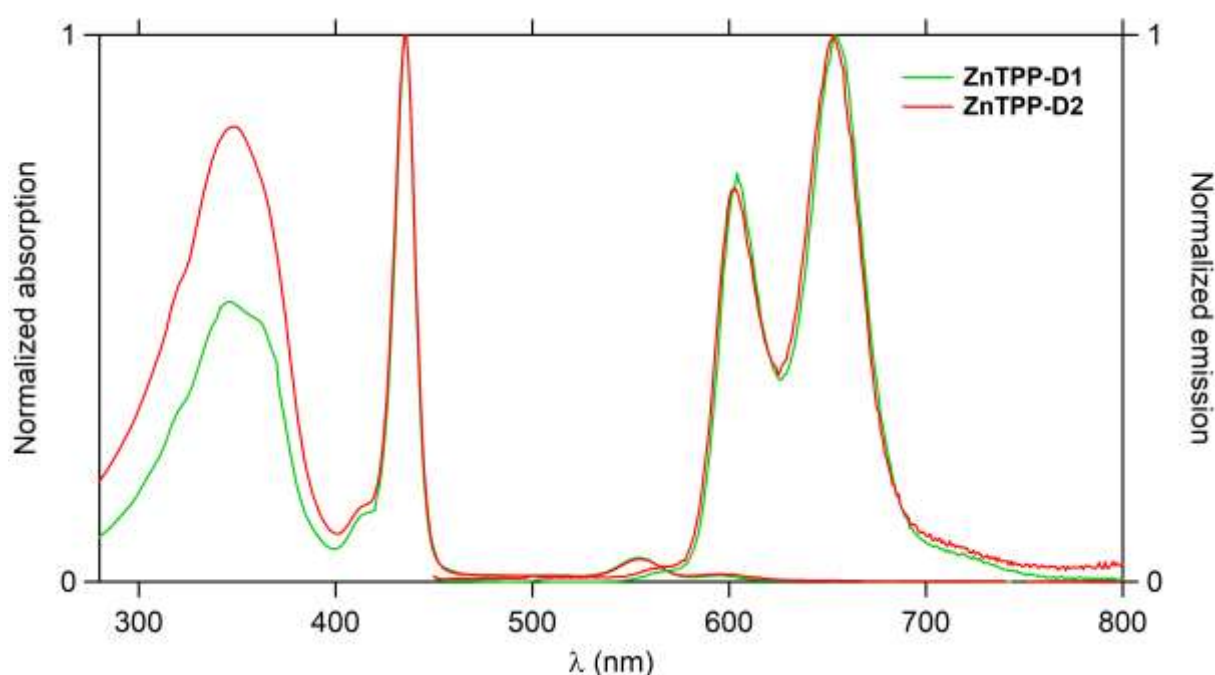


Figure 6. Absorption and emission spectra of **ZnTPP-D2** compared to **ZnTPP-D1** in CH_2Cl_2 at 20 °C.

The existence of an energy transfer (ET) process, from the peripheral 2-fluorenyl donors toward the central porphyrin acceptor core was also studied. The emission spectra were measured from 450 to 800 nm, using two excitation wavelengths; the dendron absorption (325-351 nm) and the Soret-band absorption (419-434 nm). As expected, all dendrimers show exclusively Q-band emission around 650-720 nm in both cases with is no residual dendron emission observed around 400 nm. This suggests that the peripheral fluorenyl groups transfer very efficiently their energy to porphyrin core, given that any dendron emission is totally quenched (SI Figure S14). This very efficient energy transfer most likely corresponds to a so

called “through-bond” energy transfer process (TBET).²³ Thanks to this very efficient “antenna effect”, the dendron absorption band, when intense (as in **ZnTPP-D2**), might be efficiently used for exciting these compounds.

Table 1. Photophysical properties of the new dendrimers **TPP-D1**, **ZnTPP-D1**, **ZnTPP-D2** compared to their alkynyl-bridged parents **TPP-T1** and **TPP-T2** and to **TPP** reference at 20 °C.^{7,18}

	Absorption ^a (nm)			Emission ^a Ex = Soret-band		Quantum yield ^b Φ_F (%)
	Dendron	Soret-band	Q-bands	Q(0,0)	Q(0,1)	
TPP	-	419	514, 548, 590, 649	652	719	11
ZnTPP	-	421	556, 603	603	650	3
TPP-D1	349	434	518, 552, 593, 649	652	719	13
ZnTPP-D1^c	346	436	554, 594	604	654	3
ZnTPP-D2^c	348	435	555, 595	602	652	6
TPP-T1	325	428	518, 552, 593, 649	652	719	12
TPP-T2	327, 349	428	518, 552, 593, 649	652	719	13

^a. Unless precised, experiments were achieved in toluene (HPLC level) with the UV-visible absorption region from 287 to 800 nm and emission region from 450 to 800 nm.

^b. Unless precised, experiments for fluorescence quantum yields were achieved in toluene (HPLC level) using TPP ($\Phi = 11\%$) as standard, by Soret band excitation.⁵

^c. Data obtained in CH₂Cl₂ (HPLC level).

2.3. Two-photon absorption

As these dendrimers exhibit good fluorescence properties, their intrinsic two-photon absorption cross-sections were determined by two-photon excited fluorescence (TPEF) in CH₂Cl₂. Measurements were performed with 10⁻⁴ M solutions, using a mode-locked Ti:sapphire laser delivering femtosecond pulses, following the experimental protocol described by Xu and Webb.²⁴ A fully quadratic dependence of the fluorescence intensity on the excitation power was observed for each sample at all the wavelengths of the spectra (790-920 nm), indicating that the cross-sections determined are only due to TPA. A significant increase of their TPA cross-

sections compared to that of **TPP** (12 GM at 790 nm) was observed for all porphyrins possessing fluorenyl dendrons (Table 2). Comparison between the free-base dendrimers **TPP-D1** and **TPP-T1** reveals that replacing triple bonds with double bonds in the peripheral dendrons leads to a significant increase of the TPA cross-sections.

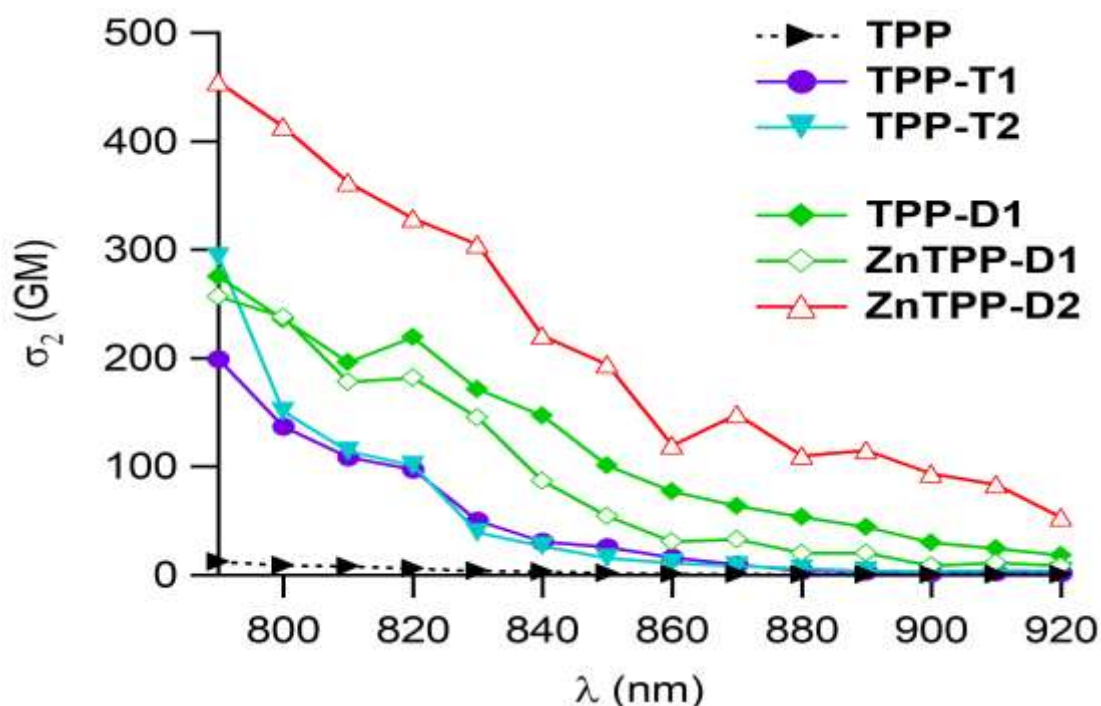


Figure 7. Two-photon excitation spectra of **D** dendrimers **TPP-D1**, **ZnTPP-D1** and **ZnTPP-D2** and comparison with related **T** compounds **TPP-T1**, **TPP-T2** and reference **TPP** in dichloromethane.

The zinc complexes of the D-type series (**ZnTPP-D1** and **ZnTPP-D2**) also exhibit high TPA cross-section ($\sigma_2 = 260$ and 450 GM, respectively) at 790 nm, **ZnTPP-D2** being the best two-photon absorber of the series of compounds presently investigated. From the comparison between **TPP-D1** and **ZnTPP-D1**, metalation by Zn(II) does not result in a significant change in cross-section at this wavelength and induces a slight decrease of σ_2 at higher wavelengths compared to the corresponding free-base porphyrin. This statement suggest that the cross-section of the missing **TPP-D2** free-base porphyrin would be similar or slightly above that of **TPP-D2** in the $790 - 920$ nm range. In this respect, the clear enhancement of σ_2 observed for **ZnTPP-D2** relative to **ZnTPP-D1** is particularly remarkable. First it reveals a more pronounced structural difference in two-photon cross-sections for the D-type series than for the T-type series, in favor of the largest dendrimers. Then it suggest that the free-base dendrimer

TPP-D2 would be a better two-photon absorber than its **TPP-T2** analogue in the investigated wavelength range.

The two-photon brightness ($\sigma_2 \cdot \Phi_F$) is a figure of merit allowing evaluating the potential of two-photon absorbers for fluorescence imaging. For the zinc complexes, any change in σ_2 is combined with a strong decrease in Φ_F leading to a decrease of this figure of merit for **ZnTPP-D1** relative to that of its free-base analogue **TPP-D1**, which is the highest of the compound presently discussed. The **ZnTPP-D2** dendrimer still exhibits an interesting two-photon brightness which is enhanced more than 20 times compared to **TPP** used as reference. Among free-base porphyrins, the two-photon brightness of **TPP-D1** is also significantly higher than that of its analog **TPP-T1**, revealing the positive impact of replacing triple bond by double bonds in the peripheral dendrons for imaging purposes.

Table 2. Two-photon absorption and brightness properties of **D** dendrimers and related and related **T** compounds in dichloromethane.

Compound	Fluorenes/ porphyrin	$\lambda_{\text{TPA}}^{\text{max}}$ (nm)	σ_2^{max} (GM) ^a	$\Phi_F \cdot \sigma_2^{\text{max}}$ (GM) ^b	Two-photon brightness enhancement factor ^c
TPP	0	790	12 ^d	1.3	1
TPP-D1	8	790	280	36	28
ZnTPP D1	8	790	260	8	6
ZnTPP-D2	16	810	450	27	20
TPP-T1	8	790	200	24	18
TPP-T2	16	790	290	38	29

^aIntrinsic TPA cross-sections measured in $10^{-4}M$ dichloromethane solutions by TPEF in the femtosecond regime; a fully quadratic dependence of the fluorescence intensity on the excitation power is observed and TPA responses are fully non-resonant.

^bMaximum two-photon brightness in dichloromethane.

^cEnhancement factor: $\Phi_F \sigma_2^{\text{max}}$ of the compound normalized to that of **TPP**.

^dData from lit.²⁵

The oxygen-photosensitizing properties of these dendrimers were also studied. Their quantum yields of singlet oxygen generation (Φ_Δ) were determined and compared to those of analogous **T** dendrimers and **TPP** used as reference (Table 3). All these dendrimers exhibit

values comprised between 0.55 and 0.64, comparable to that of reference **TPP** (0.60). Interestingly, the free-base dendrimer **TPP-D1** shows the highest value $\Phi_{\Delta} = 64\%$, whereas the two zinc complexes **ZnTPP-D1** and **ZnTPP-D2** show the lowest ones (59% and 55%, respectively). As previously noticed for T-type dendrimers^{7,18} the increase in fluorescence quantum yield of the new free-base dendrimer **TPP-D1** relative to **TPP**, is not obtained at the expense of the singlet oxygen production.

Table 3. Oxygen sensitization properties of double-bonded porphyrin dendrimers and related triple-bonded compounds.

Compound	Φ_{Δ} ^a %	$\Phi_{\Delta}\sigma_2^{\max}$ (GM) ^b	Two-photon excited oxygen sensitization enhancement factor ^c
TPP	60	7.2	1
TPP-D1	64	177	25
ZnTPP-D1	59	153	21
ZnTPP-D2	55	248	34
TPP-T1	59	118	16
TPP-T2	56	162	23

^aSinglet oxygen production quantum yield determined relative to **TPP** in dichloromethane ($\Phi_{\Delta}[\text{TPP}] = 0.60$).

^b $\Phi_{\Delta}\sigma_2^{\max}$: figure of merit of the two-photon excited singlet oxygen production in dichloromethane.

^cEnhancement factor: $\Phi_{\Delta}\sigma_2^{\max}$ of the compound normalized to that of **TPP**.

In combination with the notable increase of the TPA cross-sections of the dendrimer **TPP-D1** compared to its **TPP-T1** analogue, significant enhancements of the figure of merit for the two-photon excited oxygen sensitization ($\Phi_{\Delta}\sigma_2^{\max} = 177$ GM) can be achieved. For the zinc complexes, this value goes up to 153 GM for **ZnTPP-D1** and even to 248 GM for larger **ZnTPP-D2**. The free-base **TPP-D1** dendrimer exhibits a clearly higher enhancement factor than its T-type analogue, in relation with its higher σ_2 and its slightly increased Φ_{Δ} . This compound, easy to synthesis, appears therefore particularly promising for two-photon photodynamic therapy, and, considering its TPEF properties, also for theranostic applications provided it can be made water-soluble by proper functionalization. It should be emphasized that other porphyrin-based systems with more efficient conjugation between the sub-chromophoric

units have often been shown to exhibit higher TPA cross-sections, but these are generally accompanied by strong modifications of their other photophysical properties such as the red shift of their linear absorption range,^{10b,26} which somewhat limits their interest for theranostics. Indeed, most often these highly efficient two-photon absorbers exhibit a modest to negligible fluorescence or some interfering residual one-photon absorption above 800 nm, which leads to the loss of the 3D resolution. In contrast, the dendrimers presently reported, with a more restricted π -conjugation between the dendrons and the porphyrin core,²⁷ exhibit an improved trade-off^{7,18,28} between intrinsic TPA, fluorescence and photosensitizing properties.

3. Experimental section

3.1. General

Unless otherwise stated, all solvents used in reactions were distilled using common purification protocols,²⁹ except DMF and $i\text{-Pr}_2\text{NH}$ which were dried on molecular sieves (3 Å). All chromatographic separations were effected on silica gel (40-60 μ , 60 Å). ^1H and ^{13}C NMR spectra were recorded on BRUKER Ascend 400 and 500 at 298K. The chemical shifts are referenced to internal tetramethylsilane. High-resolution mass spectra were recorded on Bruker MicroTOF-Q II in ESI positive mode in dried solvent at CRMPO in Rennes. Reagents were purchased from commercial suppliers and used as received. Element analyses were collected on a Microanalyser Flash EA1112. UV-visible absorption and photoluminescence spectroscopy measurements for all porphyrin dendrimers in toluene were performed on Edinburg FS920 Fluorimeter (Xe900) and BIO-TEK instrument UVIKON XL spectrometer at room temperature, respectively, toluene for spectral analysis was HPLC level.

3.2. Dendron synthesis and characterization

The two dendrons **D1-PhCHO** and **D2-PhCHO**, were obtained after a multistep synthesis from the brominated 1-bromo-3,5-xylene precursors and the corresponding 2-

fluorenylaldehyde via Horner-Wadsworth-Emmons reactions followed by carbonylation using butyllithium and DMF.

1-bromo-3,5-bis(bromomethyl)benzene. Commercial 1-bromo-3,5-xylene (5.0 g, 3.67 mL, 27.02 mmol, 1 eq) was added into CH₂Cl₂ (100 mL, distilled), together with NBS (9.6 g, 54.04 mmol, 2 eq) and AIBN (220 mg, 1.35 mmol, 0.05 eq). The mixture was stirred for 30 min at room temperature, and then refluxed for 30 h. Then cooled in ice-water bath and filtered, washing residue with heptane. The solvents were evaporated and the residue was further purified by chromatography (heptane), collecting the target product (4.56 g, 49% yield) admixed with 1-bromo-3-bromomethyl-5-methylbenzene (20%) as white powder, as well as pure 1-bromo-3-bromomethyl-5-methylbenzene (3.55 g) and 1-bromo-3-(bromomethyl)-5-(dibromomethyl)benzene (1.4 g). ¹H NMR (400 MHz, CDCl₃, ppm): δ 7.47 (s, 2H), 7.34 (s, 1H), 4.41 (s, 4H).

Tetraethyl (5-bromo-1,3-phenylene)bis(methylene)diphosphonate (3). In a flask, the previously isolated [4:1] mixture of 1-bromo-3,5-bromomethyl-benzene and 1-bromo-3-bromomethyl-5-methylbenzene (2.4 g, 7 mmol, 1 eq) and P(OEt)₃ (2.4 mL, 14 mmol, 2 eq) were added respectively. The mixture was refluxed for 4 h at 140 °C. The excess of P(OEt)₃ was removed under reduced pressure. Then the title product was purified by chromatography using CH₂Cl₂ to remove other byproducts and then collected by ethyl acetate, giving a colorless oil (2.66 g, 83% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ 7.34 (s, 2H), 7.16 (s, 1H), 4.07-4.00 (m, 8H), 3.08 (d, *J* = 22.0 Hz, 4H), 1.26 (t, *J* = 7.0 Hz, 12 H).

Intermediate D1-PhBr. In a Schlenk tube, fluorenylaldehyde **1** (1.69 g, 5.51 mmol, 2.2 eq) and previously obtained **3** (1.15 g, 2.51 mmol, 1 eq) were added, then THF (100 mL, dried) was inserted. After cooling the Schlenk with an ice-water bath (0 °C), *t*-BuOK (1.20 g, 10.69 mmol, 4.4 eq) was added under Argon and the reaction was kept stirring for 1 h at 0 °C. The bath was removed, a saturated NH₄Cl solution (aq) added and the resulting solution extracted with ethyl acetate. After evaporating the solvents, it was further purified by chromatography (CH₂Cl₂/heptane [1:10]), giving **D1-PhBr** as a white powder (1.7 g, 89% yield). ¹H NMR (400

MHz, CDCl₃, ppm): δ 7.71 (d, J = 7.6 Hz, 4H), 7.61 (s, 1H), 7.60 (s, 2H), 7.52 (d, J = 8.4 Hz, 2H), 7.50 (s, 2H), 7.37-7.31 (m, 6H), 7.27 (d, J = 16.0 Hz, 2H), 7.12 (d, J = 16.4 Hz, 2H), 2.01 (t, J = 8.0 Hz, 8H), 1.15-1.06 (m, 8 H), 0.71-0.58 (m, 20H).

Dendron D1-PhCHO. In a Schlenk tube, **D1-PhBr** (720 mg, 0.94 mmol, 1 eq) was dissolved in THF (60 mL) and *n*-BuLi (0.59 mL, 0.94 mmol, 1.6 M, 1 eq) was added dropwise at -78 °C during 15 min. The reaction was stirred for additional 40 min at low temperature. Then DMF (1 mL, dried) was added and stirring was continued for 1 h at -78 °C. The bath was removed, a saturated NH₄Cl solution (aq) added and the resulting solution extracted with ethyl acetate. After evaporating the solvents, it was further purified by chromatography (CH₂Cl₂/heptane [1:5]), giving **D1-PhCHO** as a light yellow powder (580 mg, 86% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ 10.11 (s, 1H), 7.96 (s, 3H), 7.73-7.71 (m, 4H), 7.56 (d, J = 8.0 Hz, 2H), 7.53 (s, 2H), 7.40-7.23 (m, 10H), 2.02 (t, J = 8.0 Hz, 8H), 1.15-1.06 (m, 8 H), 0.71-0.56 (m, 20H). HRMS-ESI: m/z calcd for C₅₃H₅₈O: 710.44822 [M]⁺; found 710.4481.

Intermediate D2-PhBr. The synthesis is a classical procedure similar to that previously used for **D1-PhBr**. The purification was completed by chromatography (heptane/CH₂Cl₂ [10:1]), providing **D2-PhBr** as a white powder (93% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ 7.72-7.71 (m, 10H), 7.64 (s, 7H), 7.57 (d, J = 8.0 Hz, 4H), 7.55 (s, 4H), 7.37-7.31 (m, 16H), 7.30-7.18 (m, 8H), 2.03 (t, J = 8.0 Hz, 16H), 1.16-1.07 (m, 16 H), 0.72-0.57 (m, 40H).

Dendron D2-PhCHO. The synthesis is a classical procedure similar to that previously used for **D1-PhCHO**. The purification was completed by chromatography (heptane/CH₂Cl₂ [5:1]), providing **D2-PhCHO** as a yellow powder (65% yield). ¹H NMR (400 MHz, CD₂Cl₂, ppm): δ 10.15 (s, 1H), 8.05 (s, 3H), 7.76-7.73 (m, 13H), 7.62-7.59 (m, 8H), 7.42-7.29 (m, 25H), 2.06 (t, J = 8.0 Hz, 16H), 1.17-1.08 (m, 16 H), 0.72-0.54 (m, 40H). HRMS-ESI: m/z calcd for C₁₁₅H₁₂₂O: 1518.94902 [M]⁺; found 1518.9487.

3.3. Porphyrins synthesis and characterization

Reference porphyrins **TPP**, **TPP-T1** and **TPP-T2** were synthesized as described earlier by our group.^{15a-b,20} The generation G1 dendrimer **TPP-D1** was obtained under Adler Longo conditions as described earlier by our group (8% yield).¹³

ZnTPP-D1. The free-base porphyrin **TPP-D1** reacts with excess of $\text{Zn}(\text{OAc})_2$ in [3:1] mixture of CH_2Cl_2 and MeOH at 40 °C overnight. After evaporating the solvents, **ZnTPP-D1** was purified by chromatography (petroleum ether/ CH_2Cl_2 [5:1]) and after evaporation of the volatiles was obtained as a pink powder (80% yield). ^1H NMR (400 MHz, CD_2Cl_2 , ppm): δ 9.20 (s, 8H), 8.42 (s, 8H), 8.21 (s, 4H), 7.71-7.68 (m, 32H), 7.60-7.54 (m, 60H), 7.48 (d, $J = 8.4$ Hz, 8H), 7.42-7.27 (m, 80H), 7.15 (dd, $J_1 = 8.4$ Hz, $J_2 = 2.4$ Hz, 4H), 2.01 (t, $J = 7.2$ Hz, 64H), 1.12-1.01 (m, 64H), 0.65-0.51 (m, 160H). ^{13}C NMR (125 MHz, CDCl_3 , ppm): δ 151.3, 151.0, 143.0, 141.2, 140.8, 136.3, 136.1, 132.0, 130.7, 127.6, 127.1, 126.8, 125.8, 124.0, 122.8, 120.8, 119.9, 119.7, 54.9, 40.3, 25.9, 23.1, 13.8. HRMS MALDI: m/z calcd for $\text{C}_{228}\text{H}_{236}\text{N}_4\text{Zn}$: 3093.7876 $[\text{M}]^+$; found 3093.782

Dendrimer TPP-D2. The mixture of **D2-PhCHO** (250 mg, 0.16 mmol, 1 eq) and propionic acid (4 mL) was heated to 120 °C. After pyrrole (0.01 mL, 0.16 mmol, 1 eq) in propionic acid (1 mL) was added into the mixture dropwise, the reaction was kept refluxing for 5.5 h. After cooling to room temperature, MeOH was then added to the reaction mixture and the precipitate was filtered. The residue could be purified by chromatography (petroleum ether/ CH_2Cl_2 [5:1]), as a red powder (10 mg, 4% yield). ^1H NMR (400 MHz, CD_2Cl_2 , ppm): δ 9.10 (broad s, 8H), 8.50 (s, 8H), 8.20-7.10 (large signals, 188H), 2.01 (large s, 64H), 1.00 (m, 64H), 0.60-0.50 (m, 160H).

ZnTPP-D2. Previous crude mixture **TPP-D2** (10 mg, $1.6 \cdot 10^{-6}$ mol, 1 eq) reacts with excess of $\text{Zn}(\text{OAc})_2$ (3 mg, $1.6 \cdot 10^{-5}$ mol, 10 eq) in a [3:1] mixture of CH_2Cl_2 and MeOH (1 mL) at 40 °C overnight. After evaporating the solvents, the Zn complex **ZnTPP-D2**, could be isolated by chromatography (petroleum ether/ CH_2Cl_2 [5:1]), as a pink powder (60% yield). ^1H NMR (400 MHz, CD_2Cl_2 , ppm): δ 9.20 (s, 8H), 8.42 (s, 8H), 8.21 (s, 4H), 7.71-7.68 (m, 32H), 7.60-7.54 (m, 60H), 7.48 (d, $J = 8.4$ Hz, 8H), 7.42-7.27 (m, 80H), 7.15 (dd, $J_1 = 8.4$ Hz, $J_2 = 2.4$ Hz, 4H),

2.01 (t, $J = 7.2$ Hz, 64H), 1.12-1.01 (m, 64H), 0.65-0.51 (m, 160H).

3.4. Spectroscopic Measurements

All measurements have been performed with freshly-prepared air-equilibrated solutions at room temperature (298 K). UV/Vis absorption spectra were recorded on a BIO-TEK instrument UVIKON XL spectrometer or on a Jasco V-570 spectrophotometer. Fluorescence measurements were performed on dilute solutions (*ca.* 10^{-6} M, optical density < 0.1) contained in standard 1 cm quartz cuvettes using an Edinburgh Instruments (FLS920) spectrometer in photon-counting mode. Fully corrected emission spectra were obtained, for each compound, under excitation at the wavelength of the absorption maximum, with $A_{\lambda_{\text{ex}}} < 0.1$ to minimize internal absorption.

Measurements of singlet oxygen quantum yield (Φ_{Δ}). Measurements were performed on a Fluorolog-3 (Horiba Jobin Yvon), using a 450 W Xenon lamp. The emission at 1272 nm was detected using a liquid nitrogen-cooled Ge-detector model (EO-817L, North Coast Scientific Co). Singlet oxygen quantum yields Φ_{Δ} were determined in dichloromethane solutions, using tetraphenylporphyrin (**TPP**) in dichloromethane as reference solution ($\Phi_{\Delta}[\text{TPP}] = 0.60$) and were estimated from $^1\text{O}_2$ luminescence at 1272 nm.

Two-Photon Absorption Experiments. To span the 790-920 nm range, a Nd:YLF-pumped Ti:sapphire oscillator (Chameleon Ultra, Coherent) was used generating 140 fs pulses at a 80 MHz rate. The excitation power is controlled using neutral density filters of varying optical density mounted in a computer-controlled filter wheel. After five-fold expansion through two achromatic doublets, the laser beam is focused by a microscope objective (10x, NA 0.25, Olympus, Japan) into a standard 1 cm absorption cuvette containing the sample. The applied average laser power arriving at the sample is typically between 0.5 and 40 mW, leading to a time-averaged light flux in the focal volume on the order of 0.1–10 mW/mm². The fluorescence from the sample is collected in epifluorescence mode, through the microscope objective, and reflected by a dichroic mirror (Chroma Technology Corporation, USA; “blue” filter set:

675dcxr; “red” filter set: 780dxcr). This makes it possible to avoid the inner filter effects related to the high dye concentrations used (10^{-4} M) by focusing the laser near the cuvette window. Residual excitation light is removed using a barrier filter (Chroma Technology; “blue”: e650–2p, “red”: e750sp–2p). The fluorescence is coupled into a 600 μ m multimode fiber by an achromatic doublet. The fiber is connected to a compact CCD-based spectrometer (BTC112-E, B&W Tek, USA), which measures the two-photon excited emission spectrum. The emission spectra are corrected for the wavelength-dependence of the detection efficiency using correction factors established through the measurement of reference compounds having known fluorescence emission spectra. Briefly, the set-up allows for the recording of corrected fluorescence emission spectra under multiphoton excitation at variable excitation power and wavelength. TPA cross sections (σ_2) were determined from the two-photon excited fluorescence (TPEF) cross sections ($\sigma_2 \cdot \Phi_F$) and the fluorescence emission quantum yield (Φ_F). TPEF cross-sections of 10^{-4} M CH_2Cl_2 solutions were measured relative to fluorescein in 0.01 M aqueous NaOH using the well-established method described by Xu and Webb²⁵ and the appropriate solvent-related refractive index corrections.³⁰ The quadratic dependence of the fluorescence intensity on the excitation power was checked for each sample and all wavelengths.

4. Conclusions

We report here the synthesis, characterization and a photochemical study of two new zinc(II) complexes of TPP-based dendritic chromophores possessing 8 to 12 fluorenyl groups at their periphery (**ZnTPPD1** and **ZnTPP-D2**). The corresponding free-base porphyrins are analogues of related dendrimers in which we have now replaced the alkyne linkages (T series) by *E*-alkene ones (**D** series) at their periphery (Scheme 1). While **TPP-D2** could not be isolated pure, metalation of this free-base by Zn(II) provided a convenient mean to selectively access a representative of the higher generation dendrimer (**ZnTPP-D2**). Comparison with previously gathered data indicate that the optical properties of these dendrimers exhibit an obvious dependence on the dendrimer structure, *E*-alkene linkers being clearly better than 1,2-alkyne ones for enhancing the photophysical properties of interest (luminescence, 2PA cross-section

and sensitization yields) for performing 2P-PDT and fluorescence imaging. Comparison between **TPP-D1** and **ZnTPP-D1** also reveals that metalation does not drastically affect the two-photon absorption cross-section nor improve the oxygen-sensitizing efficiency of these dendrimers.

Acknowledgements

The authors acknowledge China Scholarship Council (CSC) for PhD funding (DY, LS and XZ), This project was supported by the departmental committees CD35, CD28 and CD29 of the “*Ligue contre le Cancer du Grand-Ouest*”.

References

- (1) a) Abraham, R. J.; Hawkes, G. E.; Hudson, M. F.; Smith, K. M. *J. Chem. Soc., Perkin. Trans. II* **1975**, 204-211; b) Fonda, H. N.; Gilbert, J. V.; Cormier, R. A.; Sprague, J. R.; Kamioka, K.; Connolly, J. S. *J. Phys. Chem.* **1993**, 97, 7024-7033; c) Toeibs, A.; Haeberle, N. *Justus Liebigs Ann. Chem.* **1968**, 718, 183-187.
- (2) a) Harth, E. M.; Hecht, S.; Helms, B.; Malmstrom, E. E.; Fréchet, J. M.; Hawker, C. J. *J. Am. Chem. Soc.* **2002**, 124, 3926-3938; b) Dichtel, W. R.; Serin, J. M.; Edder, C.; Fréchet, J. M. *J. Am. Chem. Soc.* **2004**, 126, 5380-5381; c) Oar, M. A.; Serin, J. M.; Fréchet, J. M. *Chem. Mater.* **2006**, 18, 3682-3692.
- (3) a) Li, B.; Li, J.; Fu, Y.; Bo, Z. *J. Am. Chem. Soc.* **2004**, 126, 3430-3431; b) Li, B.; Xu, X.; Sun, M.; Fu, Y.; Yu, G.; Liu, Y.; Bo, Z. *Macromolecules* **2006**, 39, 456-461; c) Sun, M.; Bo, Z. *Journal of Polymer Science: Part A: Polymer Chemistry* **2006**, 45, 111-124.
- (4) a) Paul-Roth, C. O.; Simonneaux, G. *Tetrahedron Lett.* **2006**, 47, 3275-3278; b) Paul-Roth, C. O.; Simonneaux, G. *C.R. Acad. Sci., Ser. IIb: Chim.* **2006**, 9, 1277-1286.
- (5) Paul-Roth, C.; Williams, G.; Letessier, J.; Simonneaux, G. *Tetrahedron Lett.* **2007**, 48, 4317-4322.
- (6) a) Drouet, S.; Paul-Roth, C.; Simonneaux, G. *Tetrahedron* **2009**, 65, 2975-2981; b) Drouet, S.; Paul-Roth, C. O. *Tetrahedron* **2009**, 65, 10693-10700; c) Merhi, A.; Drouet, S.; Kerisit, N.; Paul-Roth, C. O. *Tetrahedron* **2012**, 68, 7901-7910.

- (7) a) Yao, D.; Hugues, V.; Blanchard-Desce, M.; Mongin, O.; Paul-Roth, C. O.; Paul, F. *New J. Chem.* **2015**, 39, 7730-7733; b) Yao, D.; Zhang, X.; Mongin, O.; Paul, F.; Paul-Roth, C. O. *Chem. Eur. J.* **2016**, 22, 5583-5597.
- (8) a) L. B. Josefsen, R. W. Boyle, *Theranostics* **2012**, 2, 916-966; b) P.-C. Lo, M. S. Rodriguez-Morgade, R. K. Pandey, D. K. P. Ng, T. Torres and F. Dumoulin, *Chem. Soc. Rev.*, **2020**, 49, 1041-1056.
- (9) See for instance: a) Z. Sun, L.-P. Zhang, F. Wu, Y. Zhao, *Adv. Funct. Mater.*, **2017**, 27, 1704079; b) F. Bolze, S. Jenni, A. Sour, V. Heitz, *Chem. Commun.*, **2017**, 53, 12857-12877 and refs therein; c) M. Khurana, H. A. Collins, A. Karotki, H. L. Anderson, D. T. Cramb, B. C. Wilson, *Photochem. Photobiol.*, **2007**, 83, 1441-1448.
- (10) a) J. Schmitt, S. Jenni, A. Sour, V. Heitz, F. Bolze, A. Pallier, C. S. Bonnet, É. Tóth and B. Ventura, *Bioconjugate Chem.*, **2018**, 29, 3726-3738; b) J. R. Starkey, E. M. Pascucci, M. A. Drobizhev, A. Elliott and A. K. Rebane, *Biochim. Biophys. Acta*, **2013**, 1830, 4594-4603; c) H. A. Collins, M. Khurana, E. H. Moriyama, A. Mariampillai, E. Dahlstedt, M. Balaz, M. K. Kuimova, M. Drobizhev, V. X. D. Yang, D. Phillips, A. Rebane, B. C. Wilson, H. L. Anderson, *Nat. Photonics*, **2008**, 2, 420-424.
- (11) L. Shi, C. Nguyen, M. Daurat, A. C. Dhieb, W. Smirani, M. Blanchard-Desce, M. Gary-Bobo, O. Mongin, C. Paul-Roth, F. Paul, *Chem. Commun.* 2019, 55, 12231-12234.
- (12) Paul, F.; Lapinte, C. *Coord. Chem. Rev.*, **1998**, 178/180, 431-480.
- (13) Yao, D.; Zhang, X.; Abid, S.; Shi, L.; Blanchard-Desce, M.; Mongin, O.; Paul, F.; Paul-Roth, C. O. *New J. Chem.* **2018**, 42, 395-401.
- (14) a) Rigamonti, L.; Babgi, B.; Cifuentes, M. P.; Roberts, R. L.; Petrie, S.; Stranger, R.; Righetto, S.; Teshome, A.; Asselberghs, I.; Clays, K.; Humphrey, M. G. *Inorg. Chem.* **2009**, 48, 3562-3572; b) Yao, S.; Ahn, H.-Y.; Wang, X.; Fu, J.; Van Stryland, E. W.; Hagan, D. J.; Belfield, K. D. *J. Org. Chem.* **2010**, 75, 3965-3974; c) Mehta, G.; Sarma, P. *Tetrahedron Lett.* **2002**, 43, 9343-9346.
- (15) Hassan Omar, O.; Babudri, F.; Farinola, G. M.; Naso, F.; Operamolla, A. *Eur. J. Org. Chem.* **2011**, 529-537.
- (16) a) Adler, A.D.; Longo, F.R.; Finarelli, J. D.; Assour, J.; Korsakoff, L. *J. Org. Chem.* **1967**, 32, 476-476; b) Adler, A.D.; Longo, F. R.; Shergalis, W. *J. Am. Chem. Soc.* **1964**, 86,

3145-3149.

- (17) Lindsey, J. S.; Schreiman, I. C.; Hsu, H. C.; Kearney, P. C.; Marguerettaz, A. M. *J. Org. Chem.* **1987**, *52*, 827-836.
- (18) Yao, D.; Zhang, X.; Triadon, A.; Richy, N.; Mongin, O.; Blanchard-Desce, M.; Paul, F.; Paul-Roth, C. O. *Chem. Eur. J.* **2017**, *23*, 2635-2647.
- (19) a) M. Gouterman, *J. Mol. Spectrosc.* **1961**, *6*, 138; b) E. Austin, M. Gouterman, *Bioinorg. Chem.* **1978**, *9*, 281.
- (20) M. Gouterman, *The Porphyrins*, vol. 3, Academic press, New-York, **1978**.
- (21) Griffiths, J., *Colour and Constitution of Organic Molecules*. Acad. Press Inc.: London, **1976**.
- (22) C. O. Paul-Roth and G. Simonneaux *C.R. Acad. Sci., Ser. Iib: Chim.* **9** (**2006**) 1277.
- (23) D. Cao, L. Zhu, Z. Liu, W. Lin *J. Photochem. Photobiol. C: Photochem. Rev.* **44** (**2020**) 100371.
- (24) Xu, C.; Webb, W. W. *J. Opt. Soc. Am. B* **1996**, *13*, 481-491.
- (25) Makarov, N. S.; Drobizhev, M.; Rebane, A. *Opt. Express* **2008**, *16*, 4029-4047.
- (26) a) Drobizhev, M.; Stepanenko, Y.; Dzenis, Y.; Karotki, A.; Rebane, A.; Taylor, P. N.; Anderson, H. L. *J. Am. Chem. Soc.* **2004**, *126*, 15352-15353; b) Kim, D. Y.; Ahn, T. K.; Kwon, J. H.; Kim, D.; Ikeue, T.; Aratani, N.; Osuka, A.; Shigeiwa, M.; Maeda, S. *J. Phys. Chem. A* **2005**, *109*, 2996-2999; c) Drobizhev, M.; Stepanenko, Y.; Dzenis, Y.; Karotki, A.; Rebane, A.; Taylor, P. N.; Anderson, H. L. *J. Phys. Chem. B* **2005**, *109*, 7223-7236; d) Ahn, T. K.; Kim, K. S.; Kim, D. Y.; Noh, S. B.; Aratani, N.; Ikeda, C.; Osuka, A.; Kim, D. *J. Am. Chem. Soc.* **2006**, *128*, 1700-1704; e) Ogawa, K.; Hasegawa, H.; Inaba, Y.; Kobuke, Y.; Inouye, H.; Kanemitsu, Y.; Kohno, E.; Hirano, T.; Ogura, S.-i.; Okura, I. *J. Med. Chem.* **2006**, *49*, 2276-2283; f) Achelle, S.; Couleaud, P.; Baldeck, P.; Teulade-Fichou, M.-P.; Maillard, P. *Eur. J. Org. Chem.* **2011**, *2011*, 1271-1279; g) Hammerer, F.; Achelle, S.; Baldeck, P.; Maillard, P.; Teulade-Fichou, M.-P. *J. Phys. Chem. A* **2011**, *115*, 6503-6508; h) Pawlicki, M.; Morisue, M.; Davis, N. K. S.; McLean, D. G.; Haley, J. E.; Beuerman, E.; Drobizhev, M.; Rebane, A.; Thompson, A. L.; Pascu, S. I.; Accorsi, G.; Armaroli, N.; Anderson, H. L. *Chem. Sci.* **2012**, *3*, 1541-1547.
- (27) There is a large dihedral angle between the *meso*-aryl substituent and the macrocycle,

- which is more than 60° in the case of **TPP**, see: Silvers, S. J.; Tulinsky, A. *J. Am. Chem. Soc.* **1967**, 89, 3331-3337.
- (28) Mongin, O.; Hugues, V.; Blanchard-Desce, M.; Merhi, A.; Drouet, S.; Yao, D.; Paul-Roth, C. *Chem. Phys. Lett.* **2015**, 625, 151-156.
- (29) D. D. Perrin and W. L. F. Armarego, *Purification of Laboratory Chemicals* (3rd Edn.), Pergamon Press, Oxford, **1988**.
- (30) Werts, M. H. V.; Nerambourg, N.; Pélégry, D.; Le Grand, Y.; Blanchard-Desce, M. *Photochem. Photobiol. Sci.* **2005**, 4, 531-538.