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Evidence for anionic redox activity in a tridimensional-ordered Li-rich positive electrode -Li₂IrO₃

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Abstract

Lithium-ion battery cathode materials have relied on cationic redox reactions until the recent discovery of anionic redox activity in Li-rich layered compounds which enables capacities as high as 300 mAh g⁻¹. In the quest for new high-capacity electrodes with anionic redox, a still unanswered question was remaining regarding the importance of the structural dimensionality. The present manuscript provides an answer. We herein report on a β -Li₂IrO₃ phase which, in spite of having the Ir arranged in a tridimensional (3D) framework instead of the typical two-dimensional (2D) layers seen in other Li-rich oxides, can reversibly exchange 2.5 e⁻ per Ir, the highest value ever reported for any insertion reaction involving d-metals. We show that such a large activity results from joint reversible cationic (Mⁿ⁺) and anionic (O₂)ⁿ⁻ redox processes, the latter being visualized via complementary transmission electron microscopy and neutron diffraction experiments, and confirmed by density functional theory calculations. Moreover, β -Li₂IrO₃ presents a good cycling behaviour while showing neither cationic migration nor shearing of atomic layers as seen in 2D-layered Li-rich materials. Remarkably, the anionic redox process occurs jointly with the oxidation of Ir⁴⁺ at potentials as low as 3.4 V versus Li⁺/Li⁰, as equivalently observed in the layered α -Li₂IrO₃ polymorph. Theoretical calculations elucidate the electrochemical similarities and differences of the 3D versus 2D polymorphs in terms of structural, electronic and mechanical descriptors. Our findings free the structural dimensionality constraint and broaden the possibilities in designing high-energy-density electrodes for the next generation of Li-ion batteries.

Renewable energy sources and electric automotive transportation are popular topics in today's energy conscious society, hence placing electrochemical storage as one of the major technological challenges in this new century. Rechargeable Li-ion batteries are the technology of choice for powering electric vehicles and stand as a possible solution for grid applications provided that new advances can be made for improving the performance/cost ratio and lifetime^{1,2}. Today's research on positive electrodes is divided into two streams. One deals with polyanionic compounds that favour safety and cost at the expense of energy density³⁻⁵; and the other with chemical substitutions within the layered oxides which prioritize energy density⁶. The layered oxides are based on LiCoO₂, where Co and Li are each in 2D hexagonal arrays and Li intercalates/deintercalates via 2D migration along these sheets. Research in layered oxides has led, with the partial substitution of Co by Ni and Mn in the transition metal layer, to the presently commercialized layered oxide LiCo_{1/3}Ni_{1/3}Mn_{1/3}O₂ phase (referred to as Li-NMC) that delivers capacities ranging from 160 to 190 mAh g⁻¹ (ref. 6). Greater capacities (>280 mAh g⁻¹) were obtained by further substituting the transition metals by Li so as to produce the Li[Li_xNi_yMn_zCo_{1-x-y-z}]O₂ layered phases denoted as Li-rich NMC^{7,8}.

The staggering increase in capacity was demonstrated, via complementary electron paramagnetic resonance (EPR) and X-ray photoemission spectroscopy (XPS) measurements performed on ruthenates Li₂Ru_{1-y}Sn_yO₃ (refs 9,10), to be rooted in the anionic activity with the reversible formation of (O₂)ⁿ⁻ peroxy-like species accompanied by a massive cationic migration to the depleted Li layers. An increase in the covalence of the M-O bond, via the replacement of a 4d metal (Ru⁴⁺) by a 5d metal (Ir⁴⁺) so as to form α-Li₂IrO₃, was shown to prevent this migration¹¹. However, α-Li₂IrO₃, which enlists a structural transition on charge (from cubic close-packed O3 to hexagonal close-packed O1), presents lower cycling performances than Li₂Ru_{0.75}Sn_{0.25}O₃. Whether such a difference is nested in the existence of a structural transition or the simultaneous, rather than successive cationic and anionic redox processes for α-Li₂IrO₃ as opposed to Li₂RuO₃, remains unanswered. To address these questions, we decided to probe further the effect of modifying the crystal structure on the anionic redox reactivity.

Searching for Li-rich ordered rocksalt structures with enhanced 3D character, we became aware of the β-Li₂IrO₃ polymorph^{12,13}. Its structure exhibits an unusual topology that has recently attracted a lot of interest for magnetism¹²⁻¹⁶. This hyperhoneycomb structure, which consists in a 3D long-range ordered framework of the Ir-O bonds, differs drastically from that of layered honeycomb α-Li₂IrO₃ (Supplementary Fig. 1). It is now imperative to determine whether or not the 3D edge-sharing connectivity of the IrO₆ octahedra in β-Li₂IrO₃ can accommodate anionic distortions and thereby show the enhanced capacity associated up to now with 2D-layered Li-rich oxides

Pure β-Li₂IrO₃ powders were prepared by a high-temperature ceramic process from a mixture of IrO₂ and Li₂CO₃, as described in the Methods. The synchrotron X-ray and neutron powder

diffraction patterns were fitted using the Rietveld method, confirming the previously reported *Fddd* hyperhoneycomb structure¹³ (Fig. 1a,b and Supplementary Table 1 and Supplementary Fig. 2). β - Li_2IrO_3 is built on edge-sharing IrO_6 octahedra that form a 3D matrix in which Li ions occupy all available octahedral sites. The structure can therefore accommodate Li migration via corrugated interconnected paths, as shown in Supplementary Fig. 3.

Figure 2a shows the electrochemical performances of β - Li_2IrO_3 cycled versus Li in Swagelok cells between 2 and 4.8 V at a C/10 rate. Remarkably, all Li can be removed, leading to an ‘ IrO_3 ’ phase which can then uptake nearly 1.8 Li in the following discharge. More specifically, the first charge shows a staircase profile, with the presence of four successive plateaux at 3.45, 3.50, 4.40 and 4.55 V versus Li^+/Li^0 , that can be visualized by four peaks in the $\text{d}x/\text{d}V$ plots (Fig. 2b). Such peaks remain well defined on the subsequent discharge. This contrasts with the typical behaviour of Li-rich oxides: a staircase voltage profile on the first charge that converts to an S-shaped profile on the following discharge, which is associated with inter/intra layer cationic migrations^{9,17}. It also differs from the electrochemical behaviour of the α - Li_2IrO_3 polymorph, which shows only two oxidation plateaux that progressively transform upon cycling into S-shaped profiles¹¹. Fig. 2b also shows a minor downshift in voltage of the anodic and cathodic peaks with continued cycling up to 4.8 V, which can be avoided by limiting the cutoff voltage to 4.5 V (see Supplementary Fig. 4).

To better understand Li-driven structural changes in β - Li_2IrO_3 , in situ X-ray powder diffraction (XRD) was performed (Fig. 2c). We observe four successive biphasic regions, each of them corresponding to a well-defined plateau in the voltage-composition curve (Fig. 2c). Each inflexion point, at $x = 1.5$, $x = 1$, $x = 0.35$ and $x = 0$ in β - Li_xIrO_3 , corresponds to a single phase. The TEM analysis of these single phases reveals that the arrangement of the Ir atoms is preserved even when all Li is extracted (Supplementary Fig. 5). Refined XRD patterns (shown in Supplementary Fig. 6) demonstrate that when charging from $x = 2$ to $x = 1$, the *Fddd* space group is preserved with only a change in the lattice parameters (Fig. 2c,d). When x reaches 0.35, the symmetry of the material is lowered to the monoclinic, *C2/c* space group (note the corresponding reflection splitting in the 18.5° – 19.5° 2θ region in Fig. 2d and in Supplementary Fig. 6c and Supplementary Table 2). The fully oxidized phase Li_xIrO_3 (4.8 V, $x \sim 0$) shows a significant broadening of the peaks, most likely associated with increasing strain in the structure, and can also be indexed in the *C2/c* space group with a smaller β angle as compared to $\text{Li}_{0.35}\text{IrO}_3$ (Supplementary Fig. 6d). The feasibility of removing all Li^+ ions from β - Li_2IrO_3 , while preserving the structural framework, is somewhat spectacular as this is normally seen only in polyanionic and spinel compounds; with the main difference that two atoms out of six (that is, 33%) can be extracted from β - Li_2IrO_3 , whereas it is only one out of seven (that is, 15%) for LiFePO_4 or LiMn_2O_4 . Upon discharge, the XRD pattern of the fully charged sample converts back to that of the pristine phase via the same sequence of structure transformations. This demonstrates the

reversibility of the charge/discharge process but does not provide clues on the nature of the involved redox species.

To grasp further insights into the mechanisms of the Li-redox processes in $\beta\text{-Li}_x\text{IrO}_3$, XPS measurements were carried out on the pristine sample ($x = 2$), on ex situ oxidized samples at $x = 1.5, 1, 0.35, 0$, and on the compound that was charged to 4.8 V and discharged to 2 V versus Li^+/Li^0 . Figure 3a,b shows a shift of the Ir 4f peak towards lower energies together with the appearance of an extra O 1s peak (red peak in Fig. 3a) associated with the formation of peroxo-like $(\text{O}_2)^{n-}$ species during the charging process. During the subsequent discharge, the opposite variations are seen, indicating a reversible oxidation/reduction of both metal and oxygen upon cycling, as reported for the previously investigated Ru- and Ir-based layered Li-rich compounds^{9–11}. Moreover, Fig. 3d shows that the portion of redox that can be attributed to peroxo-like species increases monotonously until the end of charge at 4.8 V ($x = 0$), where it reaches $\sim 55\%$. Similarly, the evolution of the Ir 5d population based on the Ir valence bands (Fig. 3c) indicates that the metal oxidation takes place preferentially during the first part of charge. Overall, these results indicate that Li-deinsertion-driven cationic and anionic oxidation processes in $\beta\text{-Li}_2\text{IrO}_3$ occur conjointly, as seen in $\alpha\text{-Li}_2\text{IrO}_3$ (ref. 11).

At this stage, to explore the synergy between redox processes and structural transformations, we exploit complementary neutron powder diffraction, high-resolution scanning transmission electron microscopy (STEM) and density functional theory (DFT) calculations. An ex situ neutron powder diffraction (NPD) pattern was recorded on the sample charged to 4 V (Li_xIrO_3 with $x = 1$). The NPD pattern can be indexed with the *Fddd* orthorhombic space group (Fig. 1d), with lattice parameters reported in Table 1. Moreover, Rietveld refinements indicate that the remaining Li occupies a new tetrahedral position (Fig. 1d,e), denoted Li3 in Table 1. The structure refinement also reveals significant changes in the geometry of the IrO_6 octahedra (Fig. 1c and f). First, the average Ir–O distance reduces from 2.028 Å in the pristine structure to 1.972 Å in the charged sample, concomitantly with the decrease of the ionic radius of the Ir cations from 0.625 Å (Ir^{4+}) to 0.57 Å (Ir^{5+})¹⁸. Moreover, the shortest O–O bond in the IrO_6 octahedra is significantly shorter in the charged material (2.53 Å, Fig. 1f) than in the pristine material (2.75 Å, Fig. 1c), which is indicative of the formation of peroxo-like species on charge.

To track the deformation of the IrO_6 octahedra, we used [110] annular bright-field STEM (ABF-STEM) images of the pristine and mid-charged, $x = 1$, samples (Fig. 4), where cationic and anionic species form individual atomic columns (Supplementary Figs 7 and 8). In the image of the pristine material, the O1 and O2 atomic columns are well resolved; the image also contains a weak signal from the Li columns in the octahedral positions (see the intensity profiles). In the charged $x = 1$ material, the signal from Li in the octahedral sites is drastically reduced in line with the NPD results. The O1 positions remain clearly visible while the O2 positions become closely projected to the Ir columns. The shortening of the Ir–O2 projected distances due to the shortening of O2–O2 distances

can be seen in the [110] projection as ‘flattening’ of the IrO_6 octahedra. Together, the experimental STEM images and the NPD data unambiguously prove the formation of peroxolike species in the $x = 1$ sample, in good agreement with XPS measurements. The ABF-STEM analysis of more oxidized materials is hindered by the increasing beam sensitivity of the structures (for $x = 0.35$ image, see Supplementary Fig. 9). Nevertheless, DFT calculations reveal that the removal of one more lithium from LiIrO_3 leads to a monoclinic distortion of the structure which is in perfect agreement with the experimental data (see Supplementary Table 3) and can be explained by a significant shortening of the O1–O1 distances to reach equivalent O1–O1 and O2–O2 distances around 2.4 Å in the fully delithiated IrO_3 phase (see Supplementary Table 4).

Turning back to the cycling performances of $\text{Li}/\beta\text{-Li}_2\text{IrO}_3$ cells (Fig. 5a), we observed 15% capacity decay over the first 30 cycles when the cells are cycled between 2 V and 4.8 V versus Li^+/Li^0 , together with a 98.5% coulombic efficiency and an excellent rate capability since 80% of the initial capacity is accessible at a 4C discharge rate (Supplementary Fig. 10). To understand the origin of the capacity decay, $\text{Li}/\beta\text{-Li}_2\text{IrO}_3$ cells were cycled over different potential windows. There is only a minor decay when the cycling voltage is maintained below 4.45 V. In contrast, capacity loss is observed when the cells are cycled in the 4.45 V–4.8 V voltage domain. Electrolyte decomposition, which is enhanced at high voltage, and the poor reversibility of the structural transformation that yields ‘ IrO_3 ’ above 4.45 V, are most likely at the origin of this decay, but we cannot at present determine which one predominates.

Finally, for sake of completion, we checked the possibility of reducing further the material and, surprisingly, $\beta\text{-Li}_2\text{IrO}_3$ can reversibly uptake in a sustained way (Fig. 5b) 0.5 Li per formula unit at a potential of ~ 1.45 V versus Li^+/Li^0 via two biphasic processes (Supplementary Figs 11 and 12). Overall, the β polymorph can therefore reversibly exchange 2.5 Li. This is, to our knowledge, the first compound ever reported that cycles more than $2 e^-$ per transition metal via reversible (de)intercalation reactions.

The feasibility of preparing Li-free ‘ IrO_3 ’ electrochemically (confirmed by atomic absorption spectroscopy) fully validates our anionic redox concept. Bearing in mind the presence of Ir^{5+} and $(\text{O}_2)^{n-}$ species in this compound, we can estimate the value of n to be 3.3. Thus ‘ IrO_3 ’ can be written as ‘ $\text{Ir}^{5+}(\text{O}_3)^{5-}$ ’ to maintain charge neutrality. Such a final stage could be obtained neither with layered $\alpha\text{-Li}_2\text{IrO}_3$ nor with layered Li-rich NMC materials, mostly because they both lack the rigidity provided by the 3D hyperhoneycomb network. Similarly, this polymorph differs from $\text{Li}_{1.3}\text{Nb}_{0.3}\text{M}_{0.4}\text{O}_2$ ($\text{M} = \text{Mn}, \text{Fe}$) and $\text{Li}_{1.3}\text{Nb}_{0.43}\text{M}_{0.27}\text{O}_2$ ($\text{M} = \text{Co}, \text{Ni}$)¹⁹, which are described as cation-disordered rocksalt phases and for which the Li, Nb and M disorder precludes a long-range ordered 3D network of the metal–oxygen bonds. Moreover, because of sluggish kinetics, temperatures of 60 °C are necessary to remove one out of the 1.3 Li^+ in these phases, while all Li can be removed from $\beta\text{-Li}_2\text{IrO}_3$ at room temperature. Lastly, within this context of high-capacity materials, it is worth mentioning the recently

reported rocksalttype nanostructured ‘Li₄Mn₂O₅’ phase²⁰, whose composition and structure are not yet determined.

We have previously shown, via the study of α -Li₂IrO₃¹¹, that it is possible to suppress the typical electrochemical behaviour of Li-rich layered oxides, where a staircase voltage profile on charge is followed by an S-type profile on discharge, by increasing the covalence of the M–O bond via the use of heavier 5d metals. This finding is reinforced with the β -Li₂IrO₃ polymorph which, in contrast to the α polymorph, does not show the resurgence of the S-shape after several cycles, and remains stable after full Li removal. The elastic constants of both the α - and β -Li_xIrO₃ polymorphs were computed from DFT as a function of the Li content x (Supplementary Fig. 13). The results show that the β -phase displays a more isotropic mechanical response to stress than the α -phase, as expected from its 3D structure. Moreover, the α -Li_xIrO₃ phase at low Li content is predicted to be thermodynamically unstable with respect to shear stress, while the β -phase remains mechanically stable down to $x = 0$.

The oxidation path for β -Li₂IrO₃ counts several steps that we could ascribe to structural phase transformations. The electronic structures computed for β -Li_xIrO₃ at different Li content ($x = 2, 1, 0$) confirm the synergetic cationic and anionic redox processes at each step of the delithiation, with an equivalent contribution of Ir and O in the band lying below the Fermi level for Li₂IrO₃ and LiIrO₃ (Fig. 6a, left). Focusing on the oxygen network (from now on, O1 and O2 refer to the Fddd description), the progressive formation of peroxy-like species is clearly visible through the discretization of the O(2p)-band when going from Li₂IrO₃ to IrO₃ (see the O(2p)-band in Fig. 6a, right). The evolution of the band lying just above the Fermi level shows that O2 is mostly involved in the oxidation process from Li₂IrO₃ to LiIrO₃, while O1 is mostly involved in the oxidation process from LiIrO₃ to IrO₃. Moreover by monitoring the oxidation degree of O1 and O2 via the variation of O1–O1 and O2–O2 distances for each voltage-composition plateau (Supplementary Table 4) we could confirm, as depicted by histograms in Fig. 6b, that both oxygens are equivalently oxidized during the first (3.45 V) and the fourth processes (4.55 V) while O2 and O1 are mostly involved in the second (3.50 V) and third (4.40 V) oxidation processes, respectively. Such minor local structural changes could be at the origin of the small potential splitting in the low- and high-voltage plateaux.

In contrast, the one volt step increase at $x = 1$ could be due to the redistribution of the Li⁺ ions from the octahedral to the tetrahedral site, as observed by neutron diffraction, since a step of similar amplitude was for instance observed in spinel Li₂Mn₂O₄ upon the de-insertion of one Li when the remaining Li⁺ ions move from octahedral to tetrahedral sites. To further support this hypothesis, the electronic band structure was calculated for the β -Li₂IrO₃ and for the α -Li₂IrO₃ polymorph for sake of comparison (Supplementary Fig. 14 and Supplementary Table 5). They are nearly identical despite different crystal structures, suggesting that the topology of IrO₆–IrO₆ interactions is similar in both systems, and that equivalent electrochemical responses should be expected. This is clearly the case for the low-voltage processes, for which both phases undergo a joint cationic/anionic redox at very close

average voltages (~ 3.45 V). On the other hand, for the high-voltage process, the two polymorphs show different voltages (~ 4.45 V for β and ~ 4.2 V for α ¹¹), suggesting also the importance of the differing structural changes, namely the modification of Li sites from octahedral to tetrahedral in the β -phase and the phase transition involving a gliding of IrO_6 layers in the α - Li_2IrO_3 polymorph.

The band structures of both polymorphs are consistent with the formation of peroxo-like species at the early stage of the oxidation process. Thus, it is tempting to see the existence of a distorted IrO_6 octahedron in the pristine material as an indicator for anionic redox activity at low potential, therefore opening new directions for novel high-energy-density materials. This is certainly the case for both β - and α - Li_2IrO_3 , but further examples are needed to ensure the robustness of such a correlation. This structural feature in the pristine material could imply the presence of ligand holes in the pristine compound, not spotted by XPS, which would be neutralized by the addition of extra electrons, via classical electrochemical reduction, provided that the structure has extra sites to accommodate Li^+ ions. Along that line, the possibility to further reduce β - Li_2IrO_3 is not surprising, and fully complies with the ligand hole chemistry pioneered by J. Rouxel for chalcogenides^{21–23}.

In this report, we have demonstrated the feasibility of triggering anionic redox activity in ordered 3D compounds. Moreover, we show that a 3D structure can present some advantages such as the feasibility to remove all Li, owing to a greater stability of the structure and to achieve better capacity retention than its layered counterpart thanks to the absence of shear stresses associated with the gliding of the planes during cycling. Several interesting findings are worth mentioning. Firstly, there is a concomitant occurrence of cationic and anionic redox processes while such processes are decoupled in Li_2RuO_3 (ref. 9), with the anionic redox activity being triggered at potentials as low as 3.4 V versus Li^+/Li^0 . Secondly, the anionic redox mechanism progresses continuously on charge, irrespective of the structural transitions observed, suggesting that the voltage needed for anionic redox phenomena is closely linked to the M–O bonding and can therefore be tuned. Thirdly, we show that these redox processes do not trigger transition metal migrations as in the Ru-based layered compounds¹⁰. Lastly, the 3D β - Li_2IrO_3 compound can be both fully oxidized and reduced, a feature not present with Li-rich NMC layered oxides. These findings offer additional clues to theorists to further investigate the physics at work in these Li-rich oxides.

Overall, with the reversible uptake of 2.5 Li per transition metal, the β - Li_2IrO_3 polymorph nearly sets a record. Another unique feature of β - Li_2IrO_3 is its ability to lose all of its Li atoms, that represents 33% of the atoms in the formula unit while keeping its structural integrity. This situation has never been seen in any layered oxides. Consequently, by increasing the structural dimensionality in oxides, we now have access to a rich crystal chemistry, thereby opening wide the research field for new high-capacity electrodes relying on reversible anionic redox activity.

METHODS:

Synthesis

β -Li₂IrO₃ was prepared from IrO₂ (Sigma Aldrich, 99.9%) and Li₂CO₃ (Sigma Aldrich, 99.0%) with 20 % excess. The reactants were mixed by ball-milling 30 min, pelletized and fired at 1000°C for 86 hours under O₂ atmosphere.

Electrochemical characterization

β -Li₂IrO₃ was cycled versus Li⁺/Li in Swagelok-type cells. The pristine material was mixed with 20 wt% carbon SP by ball-milling for 20 min prior to cycling. A Whatman GF/D borosilicate glass fiber membrane was used as a separator and soaked with LP100 electrolyte. The cells were assembled in an Ar-filled glovebox and were cycled at a C/10 rate (1 Li removed in 10h) between 2 and 4.8 V versus Li⁺/Li if not otherwise specified. Samples for *ex situ* characterization were recovered inside the glovebox, washed thoroughly with anhydrous DMC and dried under vacuum.

Diffraction measurements

X-ray powder diffraction (XRD) for *in situ* measurements were done using an airtight electrochemical cell equipped with a Be window in a BRUKER D8 Advance diffractometer with Cu K α radiation ($\lambda_{K\alpha 1} = 1.54056 \text{ \AA}$, $\lambda_{K\alpha 2} = 1.54439 \text{ \AA}$). Synchrotron X-ray diffraction measurements were performed on the 11-BM beamline of the Advanced Photon Source at Argonne National Laboratory, with a wavelength of 0.41417 \AA . The powder was mixed with an appropriate amount of amorphous silica to limit absorption, and placed in a sealed quartz capillary (\varnothing 0.7 mm). Neutron powder diffraction (NPD) patterns for the pristine and 4 V samples were measured on HRPT high-resolution neutron powder diffractometer of the Swiss Spallation Neutron Source SINQ at the Paul Scherrer Institute, with a wavelength of 1.494 \AA .

Electron microscopy

Samples for the transmission electron microscopy (TEM) investigation were prepared by grinding samples in a mortar and dipping holey carbon TEM grids into the powder. To prevent exposure of the samples to air they were stored and prepared for TEM in an Ar filled glovebox. A special Gatan vacuum transfer holder was used for the analyses. High-resolution high angle annular dark field scanning transmission electron microscopy (HAADF-STEM) and annular bright field STEM (ABF-STEM) images were acquired on a probe aberration corrected FEI Titan 80-300 electron microscope at 300 kV.

X-ray photoemission spectroscopy (XPS)

XPS measurements were carried out with a Kratos Axis Ultra spectrometer, using focused monochromatic Al K α radiation ($h\nu = 1486.6$ eV). The XPS spectrometer was directly connected through a transfer chamber to an argon dry box, in order to avoid moisture/air exposure of the samples. The analyzed area of the samples was $300 \times 700 \mu\text{m}^2$. Peaks were recorded with constant pass energy of 20 eV. For the Ag 3d $_{5/2}$ line the full width at half-maximum (FWHM) was 0.58 eV under the recording conditions. The pressure in the analysis chamber was around 5×10^{-9} mbar. Short acquisition time spectra were recorded before and after each normal experiment to check that the samples did not suffer from degradation during the measurements. The binding energy scale was calibrated using the C 1s peak at 285.0 eV from the hydrocarbon contamination invariably present. To avoid any error on the calibration choice and for more precision, binding energy difference between the O1s lattice and metal core peaks was also examined. Core peaks were analyzed using a nonlinear Shirley-type background.²⁴ The peak positions and areas were optimized by a weighted least-squares fitting method using 70% Gaussian, 30% Lorentzian line shapes. Quantification was performed on the basis of Scofield's relative sensitivity factors.²⁵ The curves fit for core peaks were obtained using a minimum number of components in order to fit the experimental curves. For the valence band intensity analyses, a baseline (Shirley-type background) was chosen between 0 and 12eV approximately according to the experimental shape. The area of the band just below the Fermi level attributed to Ir5d (taking into account the differential photoionization cross sections) was normalized to the area of Ir5p, to evaluate the modification of electronic population of that band during cycling.

DFT calculations

Spin-polarized density functional theory (DFT) calculations were performed with the plane-wave density functional theory VASP (Vienna Ab initio Simulation Package) code^{26,27} using the generalized gradient approximation of Perdew–Burke–Ernzerhof (PBE) to describe electron exchange and correlation.²⁸ The rotationally invariant Dudarev method²⁹ (DFT+U) was used to correct the self-interaction error of conventional DFT for correlated d-electrons and dispersion corrections of Grimme (D3)^{30,31} were also added to better describe weak van der Waals interactions. All atom coordinates and lattice parameters were fully relaxed using conjugate gradient energy minimization until the forces acting on each atom were less than $5 \cdot 10^{-3}$ eV Å. A plane-wave cutoff of 600 eV was used to define the basis set, with well-converged k-point sampling for each compound. The elastic properties of the studied phases were investigated through the computation of the Hessian matrix of fully converged structures and we used a home-made program to plot the 3D representation of the C-matrix components, following Marnier et al.³²

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Authors Contribution

P.E. P. and J.-M.T. carried out the synthesis; P. E .P., A.J.P. and J.-M.T. did the electrochemical work; E.M. and G.R. conducted the structural analysis; M.S. and M.L.D. did the DFT calculations; D.B., A.M.A. and G.V.T. did the TEM study; D.F. collected and analysed the XPS spectra; G.R. and J.-M.T. wrote the manuscript and all authors discussed the experiments and final manuscript.

Additional Information

Supplementary information available as pdf file. CIF files of β -Li₂IrO₃ and β -LiIrO₃.

Notes

The authors declare no competing financial interest.

References

1. Tarascon, J. M. & Armand, M. Issues and challenges facing rechargeable lithium batteries. *Nature* **414**, 359–367 (2001).
2. Armand, M. & Tarascon, J. M. Building better batteries. *Nature* **451**, 652–657 (2008).
3. Padhi, A. K., Nanjundaswamy, K. S. & Goodenough, J. B. Phospho-olivines as positive-electrode materials for rechargeable lithium batteries. *Journal of the Electrochemical Society* **144**, 1188–1194 (1997).
4. Masquelier, C. & Croguennec, L. Polyanionic (Phosphates, Silicates, Sulfates) Frameworks as Electrode Materials for Rechargeable Li (or Na) Batteries. *Chemical Reviews* **113**, 6552–6591 (2013).

5. Rouse, G. & Tarascon, J. M. Sulfate-Based Polyanionic Compounds for Li-Ion Batteries: Synthesis, Crystal Chemistry, and Electrochemistry Aspects. *Chemistry of Materials* **26**, 394–406 (2014).
6. Lu, Z. H., MacNeil, D. D. & Dahn, J. R. Layered $\text{LiNi}_x\text{Co}_{1-2x}\text{Mn}_x\text{O}_2$ cathode materials for lithium-ion batteries. *Electrochem. Solid State Lett.* **4**, A200–A203 (2001).
7. Lu, Z., Beaulieu, L. Y., Donaberger, R. A., Thomas, C. L. & Dahn, J. R. Synthesis, Structure, and Electrochemical Behavior of $\text{Li} [\text{Ni}_x \text{Li}_{1/3 - 2x/3} \text{Mn}_{2/3 - x/3}] \text{O}_2$. *J. Electrochem. Soc.* **149**, A778–A791 (2002).
8. Johnson, C. S. *et al.* The significance of the Li_2MnO_3 component in ‘composite’ $x\text{Li}_2\text{MnO}_3 \cdot (1-x)\text{LiMn}_{0.5}\text{Ni}_{0.5}\text{O}_2$ electrodes. *Electrochem. Commun.* **6**, 1085–1091 (2004).
9. Sathiya, M. *et al.* Reversible anionic redox chemistry in high-capacity layered-oxide electrodes. *Nature Materials* **12**, 827–835 (2013).
10. Sathiya, M. *et al.* Origin of voltage decay in high-capacity layered oxide electrodes. *Nat Mater* **14**, 230–238 (2015).
11. McCalla, E. *et al.* Visualization of O-O peroxo-like dimers in high-capacity layered oxides for Li-ion batteries. *Science* **350**, 1516–1521 (2015).
12. Takayama, T. *et al.* Hyperhoneycomb Iridate $\beta - \text{Li}_2\text{IrO}_3$ as a Platform for Kitaev Magnetism. *Physical Review Letters* **114**, (2015).
13. Biffin, A. *et al.* Unconventional magnetic order on the hyperhoneycomb Kitaev lattice in $\beta\text{-Li}_2\text{IrO}_3$: Full solution via magnetic resonant x-ray diffraction. *Phys. Rev. B* **90**, 205116 (2014).
14. Modic, K. A. *et al.* Realization of a three-dimensional spin-anisotropic harmonic honeycomb iridate. *Nature Communications* **5**, (2014).
15. Kimchi, I., Analytis, J. G. & Vishwanath, A. Three-dimensional quantum spin liquids in models of harmonic-honeycomb iridates and phase diagram in an infinite- D approximation. *Physical Review B* **90**, (2014).
16. Kimchi, I., Coldea, R. & Vishwanath, A. Unified theory of spiral magnetism in the harmonic-honeycomb iridates α , β , and γ Li_2IrO_3 . *Physical Review B* **91**, (2015).

17. Thackeray, M. M. *et al.* Li₂MnO₃-stabilized LiMO₂ (M = Mn, Ni, Co) electrodes for lithium-ion batteries. *Journal of Materials Chemistry* **17**, 3112 (2007).
18. Shannon, R. D. Revised Effective Ionic-Radii and Systematic Studies of Interatomic Distances in Halides and Chalcogenides. *Acta Crystallographica Section A* **32**, 751–767 (1976).
19. Yabuuchi, N. *et al.* High-capacity electrode materials for rechargeable lithium batteries: Li₃NbO₄-based system with cation-disordered rocksalt structure. *Proc. Natl Acad. Sci. USA* **112**, 7650–7655 (2015).
20. Freire, M. *et al.* A new active Li–Mn–O compound for high energy density Li-ion batteries. *Nat. Mater.* **15**, 173–177 (2015).
21. Rouxel, J. Design and chemical reactivity of low dimensional solids - some soft chemistry routes to new solids. *Acs Symposium Series* **499**, 88–113 (1992).
22. Rouxel, J. Anion-cation redox competition and the formation of new compounds in highly covalent systems. *Chem.-Eur. J.* **2**, 1053–1059 (1996).
23. Rouxel, J. The Importance of Anions in redox-type Chimie Douce. *Mol. Cryst. Liq. Cryst. Sci. Technol. Sect. A-Mol. Cryst. Liq. Cryst.* **310**, 1–4 (1998).
24. Shirley, D. A. High-Resolution X-Ray Photoemission Spectrum of the Valence Bands of Gold. *Phys. Rev. B* **5**, 4709–4714 (1972).
25. Scofield, J. H. Hartree-Slater subshell photoionization cross-sections at 1254 and 1487 eV. *Journal of Electron Spectroscopy and Related Phenomena* **8**, 129–137 (1976).
26. Kresse, G. & Hafner, J. Abinitio Molecular-Dynamics for Liquid-Metals. *Physical Review B* **47**, 558–561 (1993).
27. Kresse, G. & Furthmuller, J. Efficiency of ab-initio total energy calculations for metals and semiconductors using a plane-wave basis set. *Computational Materials Science* **6**, 15–50 (1996).
28. Perdew, J. P., Burke, K. & Ernzerhof, M. Generalized gradient approximation made simple. *Physical Review Letters* **77**, 3865–3868 (1996).
29. Dudarev, S. L., Botton, G. A., Savrasov, S. Y., Humphreys, C. J. & Sutton, A. P. Electron-energy-loss spectra and the structural stability of nickel oxide: An LSDA+U study. *Phys. Rev. B* **57**, 1505–1509 (1998).

30. Grimme, S., Antony, J., Ehrlich, S. & Krieg, H. A consistent and accurate ab initio parametrization of density functional dispersion correction (DFT-D) for the 94 elements H-Pu. *The Journal of Chemical Physics* **132**, 154104 (2010).
31. Grimme, S., Ehrlich, S. & Goerigk, L. Effect of the damping function in dispersion corrected density functional theory. *Journal of Computational Chemistry* **32**, 1456–1465 (2011).
32. Marmier, A. *et al.* ELAM: A computer program for the analysis and representation of anisotropic elastic properties. *Computer Physics Communications* **181**, 2102–2115 (2010).

Table 1: Crystallographic data and atomic positions of the 4 V-charged β -LiIrO₃ determined from Rietveld refinement of neutron powder diffraction pattern. A bond valence sum analysis (BVS) (using d_0 from Ir⁴⁺) is included.

β -LiIrO ₃							
Space group $F d d d$		$R_{\text{Bragg}} = 2.64 \%$			$\chi^2 = 1.44$		
$a = 5.62149(6) \text{ \AA}$		$b = 8.88673(10) \text{ \AA}$		$c = 17.8518(2) \text{ \AA}$		$V = 891.815(17) \text{ \AA}^3$	
Atom	Wyckoff position	x/a	y/b	z/c	Occupancy	B (\AA^2)	BVS
Ir	16g	1/8	1/8	0.70899(11)	1	1.54(4)	4.555(13)
Li3	16f	1/8	0.9732(8)	1/8	1	1.84(13)	0.974(6)
O1	16e	0.8566(6)	1/8	1/8	1	1.09(5)	1.906(9)
O2	32h	0.6288(7)	0.34651(17)	0.03538(10)	1	1.42(4)	1.812(8)

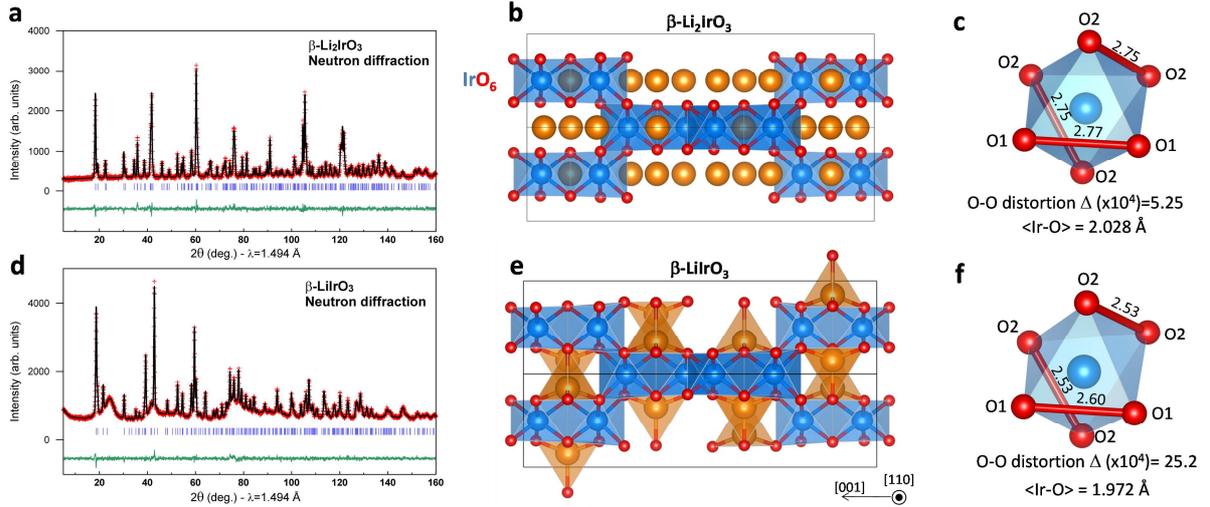


Figure 1: Structure of $\beta\text{-Li}_x\text{IrO}_3$. Rietveld refinement of the neutron patterns and resulting structures for $\beta\text{-Li}_2\text{IrO}_3$ structure (a, b, c) and 4 V-charged $\beta\text{-LiIrO}_3$ (d, e, f). In red are the experimental points, in black are the calculated patterns and in green are the differences between the experimental and calculated patterns. The vertical blue lines beneath the patterns indicate the positions of the Bragg reflections. In (d), the broad peaks at $2\theta\sim 25^\circ$, 42° , 80° are from carbon. For the structural models, Li is orange, Ir is blue, O is red. Note that Li is in octahedral position in $\beta\text{-Li}_2\text{IrO}_3$ (b) and in tetrahedral positions in 4 V-charged $\beta\text{-LiIrO}_3$ (e). Also shown is the evolution of the IrO_6 octahedron from the pristine $\beta\text{-Li}_2\text{IrO}_3$ state (c) to the 4 V LiIrO_3 compound (f). The average Ir-O distance is given (indicative of the Ir^{4+} oxidation), and we estimate the O-O distortion with the formula: $\Delta = \frac{1}{12} \sum_{i=1}^{12} \left[\frac{d_i - \langle d \rangle}{\langle d \rangle} \right]^2$ with d_i the individual O-O distances. The shortest O-O distances in the octahedral are highlighted in red with their distance (\AA) reported.

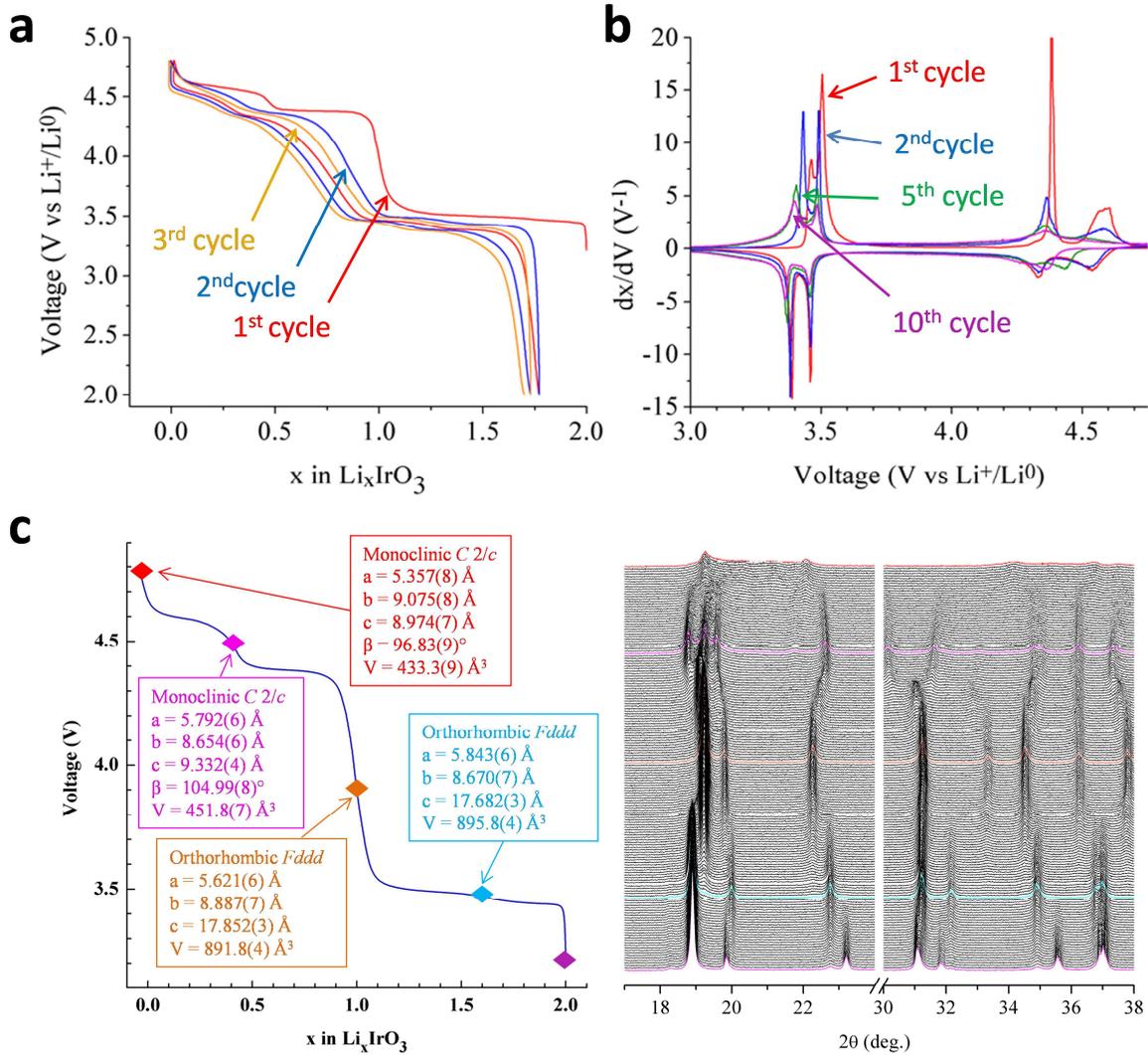


Figure 2: Electrochemical data for $\beta\text{-Li}_x\text{IrO}_3$. (a) The voltage profiles for the first 3 cycles of $\beta\text{-Li}_2\text{IrO}_3$ versus Li^+/Li^0 at C/10 rate. (b) The corresponding derivative plots dx/dV for the cycles 1 to 10. (c) The *in situ* XRD patterns of $\beta\text{-Li}_2\text{IrO}_3$ shown on the right and the corresponding charge profile on the left. The markers on the charge profile indicate the points where intermediate phases were isolated and the XRD data were refined. Space group, lattice parameters and unit cell volumes of key compositions are indicated.

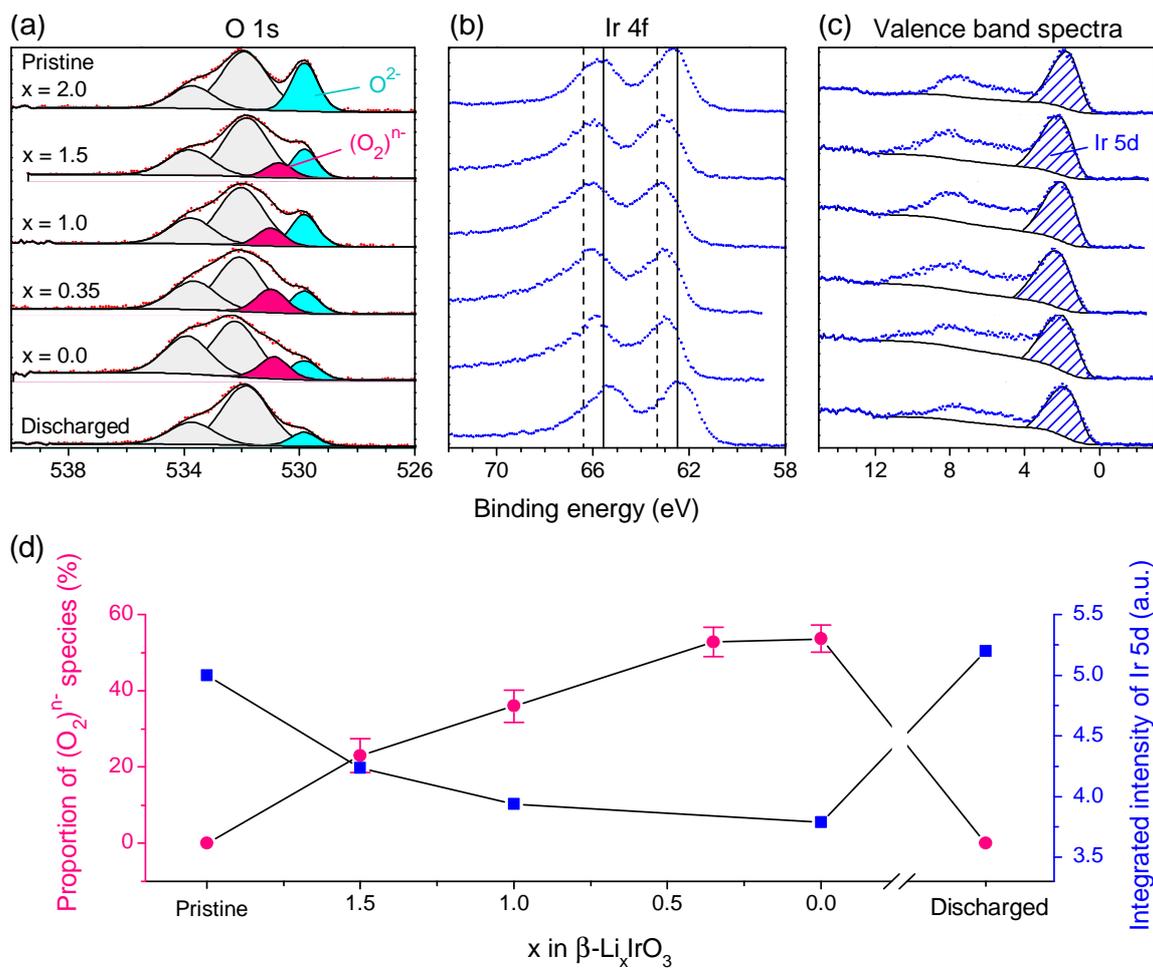


Figure 3: XPS analysis. (a) and (b) O 1s and Ir 4f XPS core spectra collected *ex situ* for $\beta\text{-Li}_x\text{IrO}_3$ at various states of charge and discharge. (c) The valence band spectra at various points in the electrochemical cycle with the hatched area representing the spectral participation of Ir 5d in this band. (d) The x dependence of the fraction of lattice oxygen attributed to peroxo-like species and the integrated intensity of the Ir 5d band, normalized to the area of Ir 5p peak.

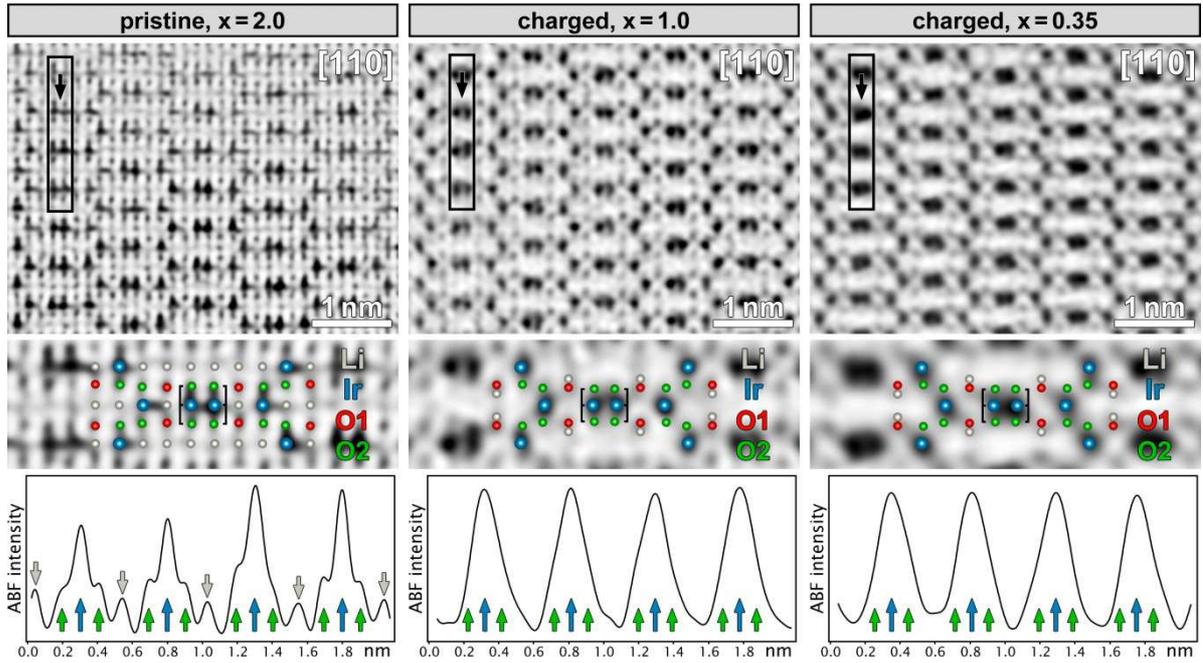


Figure 4: Local structure of $\beta\text{-Li}_x\text{IrO}_3$. [110] ABF-STEM images of the pristine $\beta\text{-Li}_2\text{IrO}_3$ and two charged Li_xIrO_3 materials with $x = 1.0$ and $x = 0.35$. The correspondence between the images and the projected structure of the materials is demonstrated on the magnified ABF-STEM in the middle ($\times 2$). Intensity profiles highlight the structure changes in the material upon charging, i.e. the vanishing of the Li signal and formation of shortened projected Ir-O2 distances. The profiles are measured in the rectangular regions indicated in the images and integrated over the width. The arrows indicate the approximate position of the atomic columns in the profiles.

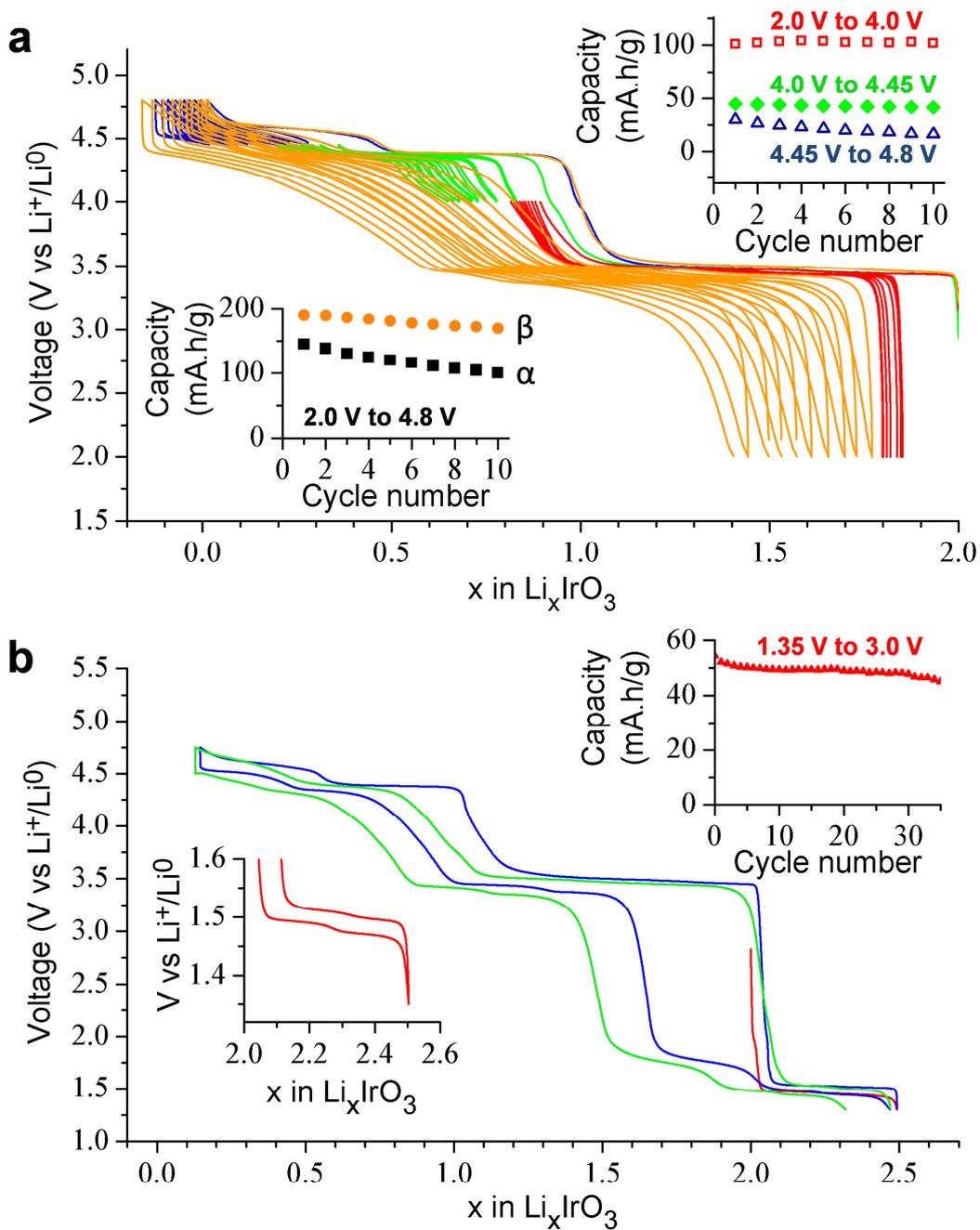


Figure 5: Cycling behavior of β - Li_xIrO_3 . (a) Voltage-composition curves for β - $\text{Li}_2\text{IrO}_3/\text{Li}$ cells cycled over different voltage ranges together with their corresponding capacity retention in the upper inset. The lower inset shows the capacity retention over the full 2 V - 4.8 V range compared to the one of the layered α polymorph cycled in the same voltage range. (b) The voltage composition profile for a similar cell started on discharge (reduction down to 1.35 V, red part of the curve together with a zoom of the low voltage region as an inset) and then cycled from 1.35 to 4.8 V (only 2 cycles are shown). The capacity retention of the low voltage plateau is shown in the upper inset.

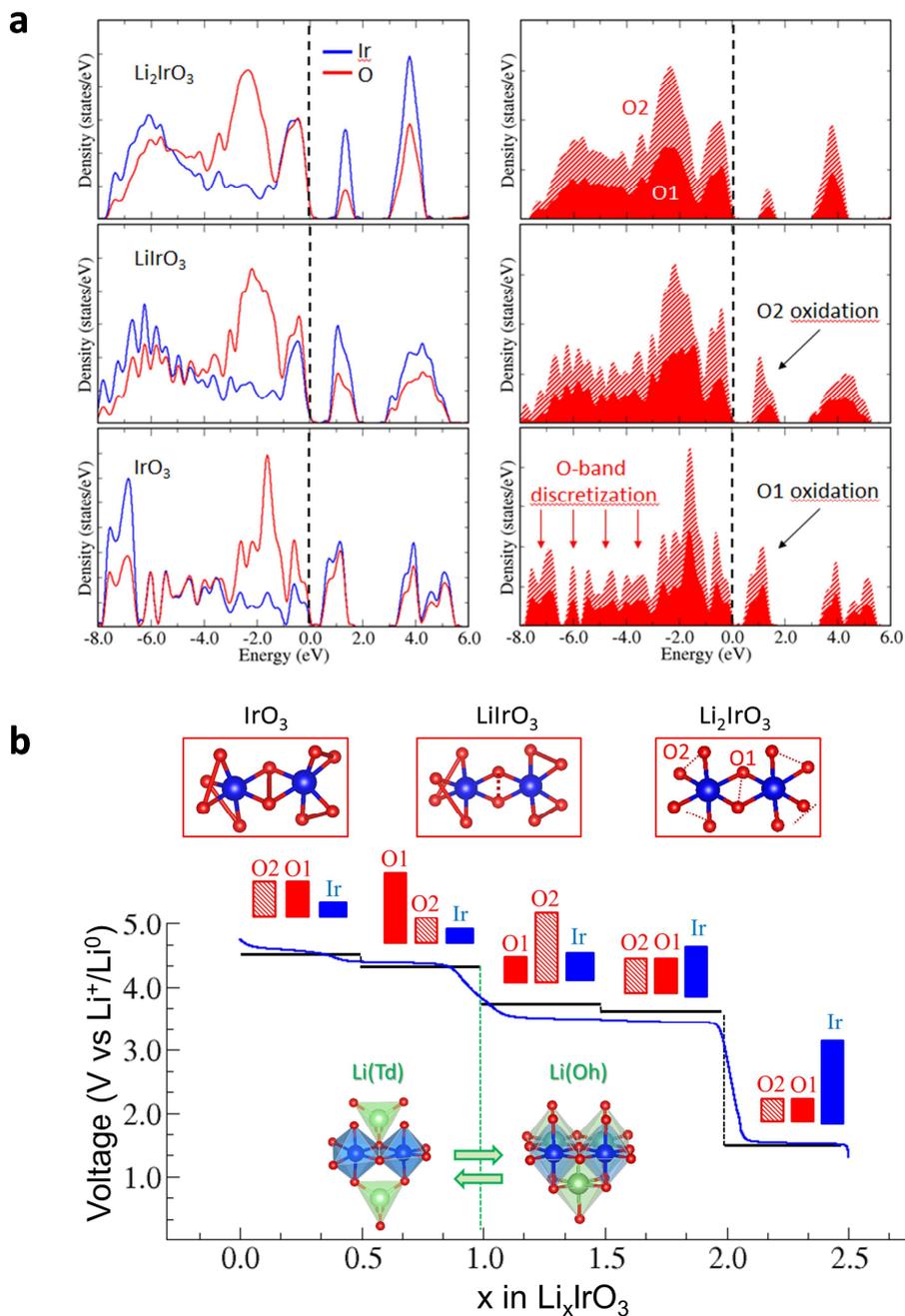


Figure 6: Electronic structures and computed voltages. (a) Density of states (DOS) projected over the Ir and O atoms (left) and over O1 and O2 atoms (right) at Li compositions Li_2IrO_3 , LiIrO_3 and IrO_3 . The Fermi level is indicated by the vertical dashed lines. Upon delithiation, the large oxygen band at $x=2$ progressively split into narrower bands corresponding to the discretization of O(2p) levels into $\sigma, \sigma^*, \pi, \pi^*$ levels of the O-O peroxo-like species. The evolution of the contribution of O1 and O2 in the band lying just above the Fermi level indicates the degree of oxidation of the different oxygens and it shows that O2 is mainly involved in the first process ($\text{Li}_2\text{IrO}_3 / \text{LiIrO}_3$) while O1 is mainly involved in the second process, in full agreement with the evolution of the O1-O1 and O2-O2 shortest distances listed in Supplementary Information, Table S4. (b) The histogram of the relative contributions of O1, O2 and Ir in each electrochemical process shown on a composition-

voltage curve where the computed average voltages (black horizontal lines) are compared to the experimental data (grey). The jump in potential from the low- to the high-voltage process is highlighted in green. The local structures of the IrO_6 dimers are given to illustrate the main distortions occurring during the first process from Li_2IrO_3 to LiIrO_3 (O2-O2 shortening) and the second process from LiIrO_3 to IrO_3 (O1-O1 shortening).