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Synthesis and properties of photoswitchable diphosphines and gold(I) complexes derived from azobenzenes

Clément Cazorla,^{a,b} Lorenzo Casimiro,^b Tanzeel Arif,^{a,b} Claire Deo,^b Nawel Goual,^a Pascal Retailleau,^a Rémi Métivier,^b Juan Xie,^{*b} Arnaud Voituriez,^{*a} Angela Marinetti^{*a} and Nicolas Bogliotti^{*b}

Diphosphines displaying azobenzene scaffolds and the corresponding bis-gold chloride complexes have been prepared and fully characterized by photophysical, spectroscopic and X-ray diffraction studies. DFT calculations provide complementary information on their electronic, structural and spectroscopic properties. Comparative investigations have been carried out on compounds featuring phosphorus functions in the *meta*- and *para*-positions, respectively, with respect to the azo functions, as well as on diphosphines with an *ortho*-tetrafluoro substituted azobenzene core. The effects of the substitution patterns on structural and spectroscopic properties are discussed.

Introduction

The chemistry of photochromic compounds including metal complexes has been the subject of strong interest over the last decade, with the aim of creating “smart” systems, able to perform a specific function or to tune their behaviour in response to light. Molecular photoswitches thus received attention as powerful tools for the external control of biological systems,^{1, 2} molecular electronic devices,³ artificial molecular machines,⁴⁻⁷ functional materials⁸ and energy storage devices,⁹ among others.¹⁰⁻¹³ Notably, photoswitchable ligands and organocatalysts have been designed for the remote control of the catalytic activity by means of non-invasive stimuli.¹⁴⁻¹⁷

Photoresponsive trivalent phosphines and their metal complexes have been only scarcely considered in this context. [For phosphines with photoresponsive dithienylethene or stiff-stilbene scaffolds, see:¹⁸⁻²³] Notably, concerning azobenzene-derived phosphines, synthetic approaches and some photophysical characterizations have been disclosed, but only sparse examples of systematic photophysical studies and uses have been reported. [Selected references on azobenzene-derived phosphines:²⁴⁻³²] Thus, having noticed that the area of photoswitchable phosphines might be somewhat underdeveloped,³³ in regard to their huge application potential in coordination chemistry and catalysis, we have initiated recently investigations in this field.

In our previous paper, we have reported on a diphosphine featuring a tetrafluoroazobenzene core, **1'**, as well as on its bimetallic gold(I) complex **1** (Fig. 1).³⁴ The catalytic activity of this complex in hydroamination reactions was found to be significantly affected by the *E/Z* isomerization of the photoswitching unit. This effect has been ascribed to an enhanced cooperativity between the two gold centers in the *Z*-form of the catalyst.

As a complement to these previous studies, we report herein the synthesis of novel azobenzene-based diphosphines and the corresponding gold(I) complexes, their structural analyses (X-ray diffraction) and detailed description of their photophysical and photochemical properties, substantiated by theoretical studies (DFT and TD-DFT). Moreover, theoretical investigations on phosphine **1'** and the gold complex **1** are also reported for comparative purpose.

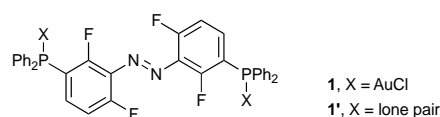


Fig. 1 Bis(diphenylphosphino)-azobenzene **1'** and the bimetallic gold complex **1**, from previous studies.³⁴

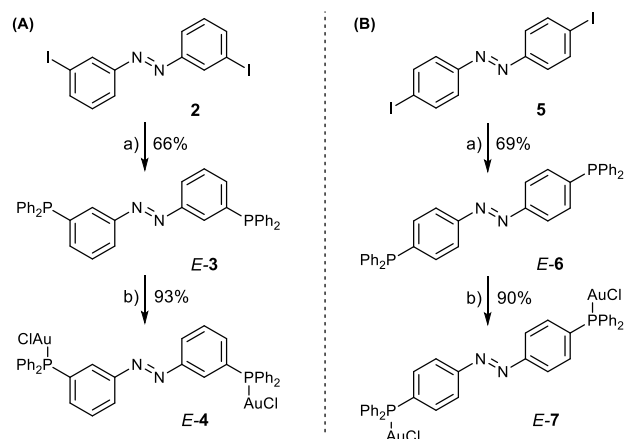
Results and discussion

Synthesis and photoisomerization studies.

Previous studies in the field of azobenzene-phosphines mainly relate to monophosphorus derivatives, while, surprisingly, the chemistry of the analogous diphosphines is largely unexplored.^{34, 35} Moreover, despite the increasing interest towards gold coordination chemistry and catalysis,³⁶⁻³⁹ only a handful of gold(I) complexes bearing either azobenzene-derived phosphines⁴⁰ or other photoswitchable phosphorus ligands have been described to date.^{18, 41, 42} In this context, we have initiated systematic investigations on azobenzene-diphosphines and their gold complexes, typified hereafter by the parent *meta*- and *para*-bis(diphenylphosphino)azobenzenes **3** and **6**. These phosphines and their Au(I)-complexes **4** and **7** have been synthesized as shown in Scheme 1. In both cases, the key step is the palladium promoted phosphination of a suitable, preformed diiodoazobenzene.²⁶

Thus, the *meta*- and *para*-diiodo-azobenzenes **2**⁴³ and **5**⁴⁴ were reacted with Ph₂PH in the presence of a catalytic amount of Pd(PPh₃)₄ and triethylamine, in toluene at 100°C. The desired azobenzene-derived diphosphines *E*-**3** and *E*-**6** were obtained in 66% and 69% yields, respectively. The gold(I)-chloride complexes *E*-**4** and *E*-**7** were obtained then in 93% and 90% yields respectively by adding ligands **3** or **6** to an Au(I) solution generated by reduction of the gold(III) precursor NaAuCl₄ with 2,2'-thiodiethanol.

Both the phosphines and the gold complexes were fully characterized and their photochemical isomerization was investigated then by spectroscopic methods (Table 1).



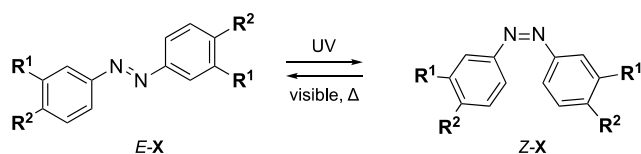
Reaction conditions:

a) Ph₂PH₂, Et₃N, cat. Pd(PPh₃)₄, toluene, 100°C, 16h.

b) NaAuCl₄(H₂O)₂, 2,2'-thiodiethanol, H₂O, then **3** or **6**, CHCl₃, 0°C

Scheme 1 Synthesis of the 3,3'- and 4,4'-bis(diphenylphosphino)azobenzenes **E-3** and **E-6**, and their gold complexes **E-4** and **E-7**.

Table 1 ³¹P NMR chemical shifts and UV-vis absorption max for the azobenzene-phosphines and gold(I) complexes. Calculated max for pure Z compounds



X	R ¹	R ²	³¹ P NMR ^a E-X	UV-vis ^b E-X		UV-vis ^b Z-X	
				π-π* [nm]	n-π* [nm]	π-π* [nm]	n-π* [nm]
3	PPh ₂	H	-4.9	320	447	257	434
4	PPh ₂ (AuCl)	H	33.6	320	449	270 ^c	433
6	H	PPh ₂	-4.6	365	470 ^c	320 ^c	440
7	H	PPh ₂ (AuCl)	33.3	333	456	262 ^c	433

^a chemical shift in CDCl₃; ^b λ_{max} in MeCN; ^c shoulder

The photophysical and photochemical characterization of all the compounds was carried out in MeCN. However, the scarce solubility in this solvent did not allowed a reliable absorption coefficient determination. The absorption spectrum of the *meta*-bis(diphenylphosphino)-azobenzene **E-3** (Fig. 2) shows an intense absorption band at 320 nm, corresponding to the π-π* electronic transition, and a weaker band at 447 nm, corresponding to the forbidden n-π* transition. The energy of these absorption maxima is in line with those of the unsubstituted azobenzene.⁴⁵ In the gold complex **E-4**, the position of the two bands is substantially unchanged (λ_{max} = 320 nm and 449 nm), although the structure of the π-π* band becomes more pronounced. Contrariwise, if the phosphines are placed in *para*-position, the electronic effect exerted on the absorption spectrum is stronger. For the *para*-ligand **E-6**, the π-π* band is centered at 365 nm – redshifted of 45 nm with respect to **E-3** – and partially overlapped to the n-π* band. In

the corresponding gold complex **E-7**, the redshift is less intense and the π-π* and n-π* bands are better resolved, being observed at 333 nm and 456 nm, respectively.

The optical features of **E-6** would suggest that, when the functionalization is in *para*-position, the electron-donating phosphorus atoms are involved in the conjugation, as it happens for amino-azobenzene.⁴⁵ This effect is less intense in **E-7**, where the lone pair is involved in the complexation of the metal. Conversely, when the functionalization is in *meta*-position, the phosphorus atoms are less involved in the conjugation and the resulting spectrum is comparable with azobenzene-type compounds and insensitive to the complexation.

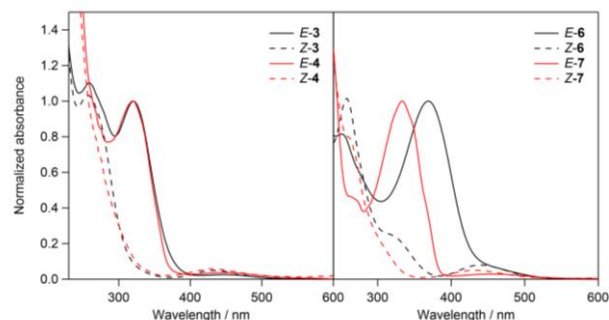


Fig. 2 Left: absorption spectrum in MeCN of the *E*- (full lines) and *Z*- (dashed lines, extrapolated spectra) isomers of **3** (black, left) and **4** (red, left) and of **6** (black, right) and **7** (red, right).

Upon UV irradiation (313 nm, 334 nm or 365 nm), under inert atmosphere, all the compounds displayed absorbance variations consistent to a *E*-to-*Z* photoisomerization (Fig. 3): the π-π* band decreased, while the n-π* increased, until a photostationary state (PSS) was achieved. In the case of the gold complexes, the n-π* band was also blueshifted of few nanometres with respect to the *E* isomer. In the same way, visible irradiation (436 nm) on solutions previously irradiated in the UV promoted *Z*-to-*E* photoisomerization (Fig. S11).

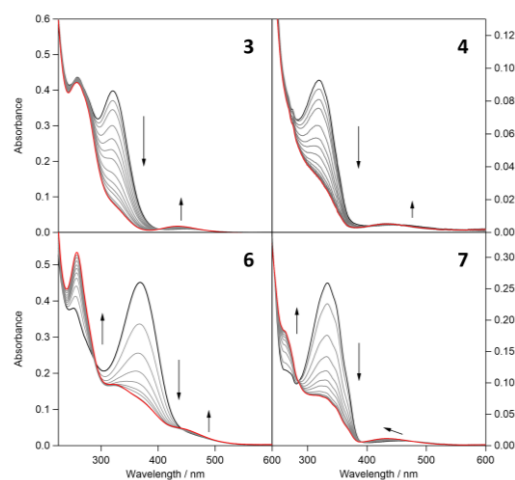


Fig. 3 Absorption evolutions in MeCN upon irradiation at 365 nm of **3** (top left), **4** (top right), **6** (bottom left), **7** (bottom right).

By applying the Fischer method⁴⁶ to the spectra of photostationary states obtained upon irradiation at different wavelengths (Fig. S12), we were able to extract the spectrum of the pure Z form (Fig. 2) and to calculate consequently the E/Z ratio at each PSS. Generally, if UV light is employed, all the *E*-compounds reach good photoconversion rates, with up to 87% conversion observed for *E*-**3** under irradiation at 334 nm (Table 2). If visible light is employed, the PSSs are more shifted toward the *E* form. Interestingly, upon irradiation of an *E/Z*-**7** mixture at 436 nm, an *E:Z* ratio of 96:4 was reached, meaning that visible light can trigger an almost complete reset to the *E* form of the gold complex **7** (Table 2).

Through fatigue resistance experiments, we proved that the photoswitching between two different PSSs is completely reversible and can be repeated up to 10 cycles without any evidence of photodegradation (Fig. S13).

We investigated the thermal isomerization kinetics at 60°C after exhaustive irradiation of the *E*-forms of **3**, **4**, **6** and **7** at 365 nm (see Fig. 4, S14 and Table 3). In line with the optical features, the resulting kinetic constants confirmed that for the *meta*-derivatives, where the electron-donating effect of the phosphorus atom is less prominent, the thermal isomerization is slower, with rate constants in the order of 10^{-5} s^{-1} , corresponding to half-lives $t_{1/2}$ ranging from 268 to 326 minutes. For the *para*-derivatives, the thermal isomerization is about 10-fold faster, in the order of 10^{-4} s^{-1} ($t_{1/2} = 54\text{-}98$ minutes).^{47,48}

Table 2. UV-vis photoconversions of azobenzene-phosphines and gold complexes.

	E:Z 313 nm ^a	E:Z 334 nm ^a	E:Z 365 nm ^a	E:Z 436 nm ^a
3	35:65	13:87	13:87	81:19
4	44:56	18:82	30:70	83:17
6	75:25	60:40	18:82	72:28
7	74:26	60:40	25:75	96:4

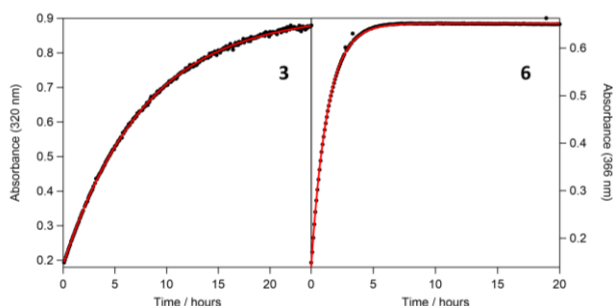


Fig. 4 Absorbance evolution in the dark at 60°C in MeCN (black dots) for **3** (left, 320 nm) and **6** (right, 366 nm), and fitting (red line), after exhaustive irradiation at 365 nm.

The half-lives of the trivalent phosphine *Z*-**3** and its gold(I) complex *Z*-**4** are of more than 50-fold lower than those of the corresponding *o*-difluorinated phosphine **1'** and gold complex **1** (Fig. 1), that were measured at 219 h and 278 h respectively.³⁴ This effect is fully consistent with literature data that have demonstrated the increased thermal stability of *o*-fluorine-substituted *Z*-azobenzenes.⁴⁹⁻⁵² The effect has been ascribed, in polar solvents, to the increased dipole moment of the *Z*-isomers, induced by the fluoride substituents.⁵³

Table 3. Thermal isomerization parameters of azobenzene-phosphines and gold complexes.

	k^a	$t_{1/2}^b$
3	$3.54 \times 10^{-5} \text{ s}^{-1}$	326 min
4	$4.31 \times 10^{-5} \text{ s}^{-1}$	268 min
6	$2.15 \times 10^{-4} \text{ s}^{-1}$	54 min
7	$1.18 \times 10^{-4} \text{ s}^{-1}$	98 min
1'	---	219 h ^c
1	---	278 h ^c

^a Kinetic constant for the thermal isomerization at 60 °C in MeCN; ^b Half-life at 60 °C in MeCN; ^c from ref. ³⁴

X-ray diffraction and DFT studies.

Single crystals of the bimetallic gold(I) complexes *E*-**4** and *E*-**7**, suitable for X-ray diffraction studies, were obtained by slow evaporation of a chloroform solution. ORTEP drawings for the two complexes are displayed in Fig. 5. Selected bond lengths and angles are given in Table 4. For comparison, X-ray data for the fluorinated azobenzene phosphine complexes *E/Z*-**1** are also given.³⁴

Both diphosphine complexes **4** and **7** adopt extended conformations, with the gold centers on opposite faces of the plane defined by the azobenzene core. Both complexes exhibit planar arrangements of the azo unit, as revealed by the 180.0° values of the dihedral angle Ψ (C(8)-N(10)-N(11)-C(12) angles). Bond distances are similar for the two compounds: Au-Cl, 2.277 and 2.279 Å, Au-P, 2.230 and 2.227 Å; N-C, 1.423 and 1.418 Å for *E*-**4** and *E*-**7** respectively. In the *m*-substituted complex *E*-**4** the N-N, bond (1.255 Å) is slightly longer than in *E*-**7** (1.238 Å). For comparison, the N-N bond length of the non-substituted *trans*-azobenzene measures 1.249 Å.⁵⁴

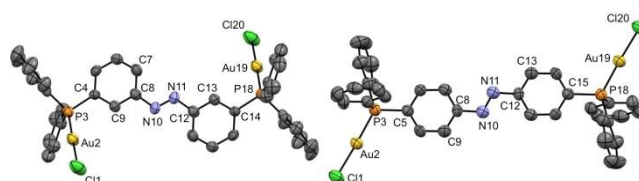


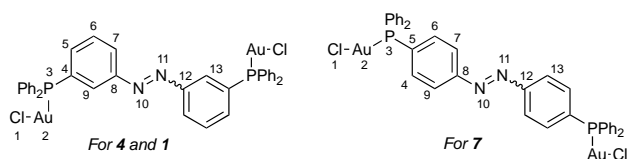
Fig. 5 ORTEP drawings for complexes *E*-**4** (left) and *E*-**7** (right), thermal ellipsoids drawn at the 30% probability level.

Table 4. Selected bond lengths (Å) and angles (deg) for the gold complexes *E-4*, *E-7*, *E-1* and *Z-1*, from X-ray crystal structures.

	<i>E-4</i>	<i>E-7</i>	<i>E-1</i> ³⁴	<i>Z-1</i> ³⁴
Cl(1)–Au(2)	2.277(2)	2.279(2)	2.276(8)	2.280(2)
Cl(20)–Au(19)	2.277(2)	2.279(2)	2.276(8)	2.269(2)
Au(2)–P(3)	2.230(2)	2.227(2)	2.226(7)	2.228(1)
Au(19)–P(18)	2.230(2)	2.227(2)	2.226(7)	2.223(1)
P(3)–C(4)	1.815(7)	-	1.82(2)	1.824(6)
P(18)–C(14)	1.815(7)	-	1.82(2)	1.824(6)
P(3)–C(5)	-	1.832(5)	-	-
P(18)–C(15)	-	1.832(5)	-	-
N(10)–C(8)	1.423(9)	1.418(7)	1.47(3)	1.447(7)
N(11)–C(12)	1.423(9)	1.418(7)	1.47(3)	1.436(7)
N(10)–N(11)	1.255(9)	1.238(6)	1.26(3)	1.238(8)
Ψ C(8)–N(10)–N(11)–C(12) ^b	180.0(6)	180.0(4)	180(2)	2.9(8)
C(9)–C(8)–C(12)–C(13)	180.0	180.0	180	29.5
Φ C(7)–C(8)–N(10)–N(11)	14.1	4.6(8)	42(3) ^c	60.8(8)

^a Values from one of the two independent molecules present in the unit cell. ^b Absolute value. ^c Φ C(9)–C(8)–N(10)–N(11).

Atom numbering for compounds in the Table:



The *o*-fluorine substituents in **1** do not affect substantially bond parameters. However, it appears clearly that they induce a rotation of the aryl rings out of the C=N=N-C plane (Fig. 6). In the solid state, the dihedral angle Φ between the phenyl rings and the azobenzene plane increases indeed from 14.1° to 42° moving from non-fluorinated *E-4* to the fluorinated analogue *E-1*. Rotation of the aryl rings out of the plane of the diazo bond had been reported notably for the parent tetra-*o*-fluoroazobenzene.⁵⁵ Its optimized geometry from DFT calculations showed a dihedral angle Φ = 27.7°, which is broadly consistent with the experimental value observed for *E-1*.

The *p*-diphenylphosphino-substituted compound *E-7* displays an almost planar arrangement, with a dihedral angle Φ = 4.6°.

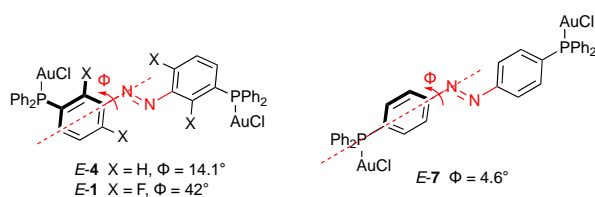


Fig. 6 Structural features of complexes *E-4* and *E-1* and *E-7*: dihedral angles Φ between the C=N=N-C plane and the aryl planes.

These structural data are indicative of the preference of non-substituted azobenzenes and *o*-substituted azobenzenes for *planar* and *non-planar arrangements*, respectively. However, the absolute Φ values should be taken with caution since azobenzenes display conformational flexibility in solution and a high degree of variation of dihedral angles in the solid state.^{54, 55}

To get insights into the electronic features of the *E/Z*-diphosphines and gold complexes **3**, **4**, **6**, **7**, **1'** and **1**, geometry optimization were performed at the B3LYP level using the 6-311G(d,p) (for **3**, **4**, **6** and **7**) and the 6-311++G (for **1** and **1'**) basis sets (LANL2DZ for Au atom), taking into account the effect of solvent (MeCN) with C-PCM model.⁵⁶ Views of the optimized geometries of *Z-4* and *Z-7* are given in Fig. 7 as representative examples. Other data are reported in the SI section. For *Z-4* and *Z-7*, the calculated bond distances and angles do not show significant changes with respect to the corresponding *E*-isomers, except for dihedral angle values. Notably, as expected, in the *Z*-forms the two aryl rings are twisted out of the azobenzene plane so as to minimize the mutual steric hindrance (Φ = 53.3 and 53.4° for *Z-4* and *Z-7* respectively). Other relevant dihedral angle values are: Ψ C(8)–N(10)–N(11)–C(12) = 8.5 and 9.2, for *Z-4* and *Z-7* respectively, and C(9)–C(8)–C(12)–C(13) = 95.2 and 95.8 (for *Z-4* and *Z-7*).

The almost invariant main bond parameters emphasize the ability of azobenzene-derived diphosphine ligands to photocontrol the relative orientation of the two metal centers in their complexes, without affecting the other main bond parameters.

The electronic transitions relevant to both *E*- and *Z*-ligands **3** and **6** and the Au(I) complexes **4** and **7** were calculated by TD-DFT at the CAM-B3LYP level using the 6-311+G(d,p) basis set (LANL2DZ for Au) and C-PCM model for MeCN.⁵⁶

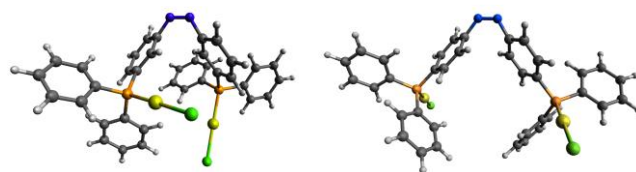


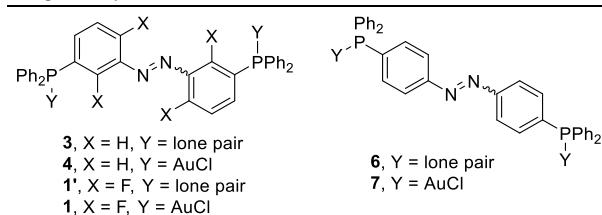
Fig. 7 DFT-optimized structures of complexes *Z-4* and *Z-7* at the C-PCM-B3LYP/6-311G(d,p)[H,C,N,P,Cl]/LANL2DZ[Au] level in MeCN.

The involvement of multiple orbitals precluded straightforward visualization of electronic excitation, thereby leading us to generate natural transition orbitals (NTOs) for a qualitative description.⁵⁷

The calculated characteristic UV and visible bands of the *E* and *Z* isomers showed good agreement with the experimental values (Table 5). Indeed, for ligand *E-3* and the Au(I) complexes *E-4* and *E-7*, calculations predict absorption bands between 315 and 339 nm, in keeping with the experimental spectra

measured in MeCN (320–333 nm). These transitions involve orbitals localized on the azobenzene core and exhibit a dominant $\pi \rightarrow \pi^*$ character (Fig. 8a and SI11). Interestingly, the redshift to 365 nm observed in the case of *E*-**6** (bearing a *diphenylphosphino*-substituent in *para* position to the N=N bond), is accurately predicted by calculations (363 nm). It is found to originate from the *extended overlap of the azobenzene π system with orbitals of the P atom* (Fig. 8b).

Table 5. Characteristic absorption wavelengths for the *E/Z* isomers of the diphosphines and gold complexes



	$S_0 \rightarrow S_2$ [nm]		$S_0 \rightarrow S_1$ [nm]	
	exp. ^a	calc. ^b	exp. ^a	calc. ^b
<i>E</i> - 3	320	330 ^c	447	454 ^{d,e}
<i>Z</i> - 3	257	286	434	456 ^d
<i>E</i> - 4	320	315 ^c	449	457 ^{d,e}
<i>Z</i> - 4	270 ^f	274	433	451 ^d
<i>E</i> - 6	365	363 ^c	470 ^f	458 ^d
<i>Z</i> - 6	320 ^f	302	440	465 ^d
<i>E</i> - 7	333	339 ^c	456	469 ^{d,e}
<i>Z</i> - 7	262 ^f	282	433	459 ^d
<i>E</i> - 1'	304	335 ^c	455	508 ^{d,e}
<i>Z</i> - 1'	nd	304	418	441 ^d
<i>E</i> - 1	310 ^g	325 ^c	456 ^g	509 ^d
<i>Z</i> - 1	293 ^{f,g}	283	422 ^g	441 ^d

^a Experimental values in MeCN solution. ^b calculated values in MeCN using the C-PCM-TD-CAM-B3LYP/6-311+G(d,p) [H,C,N,P,Cl]/LANL2DZ[Au]/C-PCM-B3LYP/6-311G(d,p)[H,C,N,P,Cl]/LANL2DZ[Au] method for *E/Z*-**3/4/6/7**, the C-PCM-TD-CAM-B3LYP/6-311+G(d,p)[H,C,N,F,P,Cl]/LANL2DZ[Au]/C-PCM-B3LYP/6311++G[H,C,N,F,P,Cl]/LANL2DZ[Au] method for *E/Z*-**1/1'**. ^c Shows dominant $\pi \rightarrow \pi^*$ character. ^d Shows dominant $n \rightarrow \pi^*$ character. ^e Shows oscillator strength close to 0. ^f Shoulder. ^g From ref ³⁴. nd: not determined.

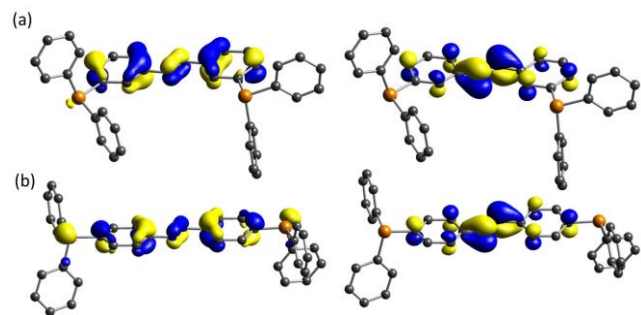


Fig. 8. Dominant natural transition orbitals pairs from holes (left) to particles (right) for electronic transition of (a) *E*-**3** at 330 nm and (b) *E*-**6** at 363 nm.

The $S_0 \rightarrow S_1$ transitions (showing a dominant $n \rightarrow \pi^*$ character) for both phosphine *E*-**3** and complex *E*-**4** (447 and 449 nm, respectively) are blue-shifted compared to the fluoro-substituted analogues *E*-**1'** and *E*-**1** (455 and 456 nm, respectively). The calculated shorter N–N bond lengths in *E*-**3/E**-**4** (1.254 and 1.253 Å, respectively) relative to *E*-**1'/E**-**1** (1.281 and 1.280 Å, respectively), as well as the lower Φ angle in *E*-**3/E**-**4** (0.9° and 1.2°, respectively) as compared to *E*-**1'/E**-**1** (3.5° and 6.7°, respectively, see Table SI1 and Fig. SI 6,7 and 9,10) could be responsible for the observed hypsochromic shift.⁵⁵

The *Z*-isomers of **3**, **4**, **6** and **7** exhibit characteristic absorption bands in the visible region between 434 and 470 nm (calculated in the range 451–465 nm) which are clearly assigned to $n \rightarrow \pi^*$ transitions located at the N=N moiety of the azobenzene core (Fig. 9 and SI12). The absorption bands in phosphine *Z*-**3** and gold complex *Z*-**4** (434 and 433 nm, respectively) appear at longer wavelengths than in the fluoro-substituted analogues *Z*-**1'** (Fig. SI 5) and *Z*-**1** (418 and 422 nm, respectively). A similar hypsochromic shift has been documented previously for the *Z*-isomers of fluorine-substituted azobenzenes vs non-substituted azobenzenes.⁵⁸ It should be related to the increased withdrawal of electron density around the diazene unit.

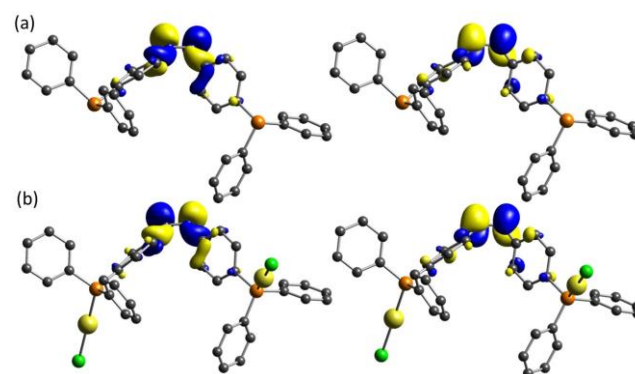


Fig. 9. Dominant natural transition orbitals pairs from holes (left) to particles (right) for electronic transition of (a) *Z*-**6** at 465 nm and (b) *Z*-**7** at 459 nm.

As expected, the extent of $n \rightarrow \pi^*$ band separation between *E* and *Z* isomers is significantly increased in the presence of fluorine atoms, as revealed by $\Delta\lambda$ ($n \rightarrow \pi^*$, *E/Z*) values raising from 13 and 16 nm (for **3** and **4**, respectively) to 37 and 34 nm (for **1'** and **1**, respectively)

Conclusions

Two novel azobenzene-based diphosphines **3** and **6** displaying phosphorus functions in *meta*- and *para*-positions respectively, have been prepared and their gold(I) complexes have been characterized by X-ray diffraction studies. Absorption spectroscopy and computational methods (DFT and TD-DFT) have been performed to characterize the electronic properties of these compounds. These studies show that the electronic features of azobenzene are almost unchanged when diphenylphosphino functions are introduced on its azo

positions, while bathochromic shifts of the $\pi \rightarrow \pi^*$ absorption bands are induced by phosphorus functions in *para*-positions. The *E-Z* photochemistry of the diphosphines and their complexes has been elucidated via UV-vis absorption spectroscopy. The absorption spectra of the *meta*-derivatives resulted to be similar to the underivatized azobenzene and rather insensitive to the complexation. Conversely, the enhanced electronic effect exerted by the phosphines in *para*-position confers to the azobenzene diphosphine **6** an amino-azobenzene character, which becomes less intense in its Au(I)-complex **7** due to the complexation. All the compounds showed a good *E* \rightarrow *Z* photoconversion upon UV irradiation and afforded almost complete recovery of the *E*-isomer upon visible irradiation. The thermal isomerization half-lives resulted to be below 100 minutes for the *para*-derivatives and higher (270-330 minutes) for the *meta*-derivatives.

Compounds **3** and **4** have been compared to the parent *o*-fluorinated *m*-azobenzene diphosphine **1'** and gold complex **1**, through both experimental and theoretical studies. The results demonstrate the beneficial effect of di-*ortho*-fluorination, that leads to an increased separation of the $n \rightarrow \pi^*$ bands between *E* and *Z*-isomers, to a better photoconversion extent, as well as to an increased thermostability of the *Z*-forms.

Overall, the investigations above afford preliminary information on the structural and photophysical properties of azobenzene-diphosphines, as a function of their substitution patterns. Such knowledge is an essential prerequisite to develop further applications of azobenzene-diphosphines and their metal complexes in various fields, including photoswitch catalysis.

Experimental section

General

All reactions requiring oxygen free conditions were carried out under an argon atmosphere. Toluene and THF were dried before use. Commercially available starting materials were used without further purification. Purification of the products was carried out by flash chromatography using silica gel (35-70 μ M). ^1H , ^{31}P and ^{13}C NMR spectra were recorded in CDCl_3 on Bruker AV500 or AV300 spectrometers. Chemical shifts are reported in ppm and referenced to residual CHCl_3 at δ 7.26 for ^1H NMR spectra, and at 77.16 for ^{13}C NMR spectra. Low and high resolution mass spectra [MS(TOF Cl+)] were recorded on a LCT Waters equipment. Crystals suitable for X-ray analysis were obtained by slow evaporation of a chloroform solution.

General procedure for the synthesis of phosphines **E-3** and **E-6**.

A flame dried Schlenk was charged with the diiodide **2**⁴³ (or **5**⁴⁴) (0.23 mmol, 100 mg), $\text{Pd}(\text{PPh}_3)_4$ (5 mol %, 13 mg) and toluene (1.2 mL). To the resulting solution, triethylamine (0.69 mmol, 96 μ L) and diphenylphosphine (0.51 mmol, 88 μ L) were added. The mixture was stirred at 100°C during 16h. After cooling to room temperature water was added, the organic layer was separated and the aqueous layer was extracted with dichloromethane. The combined organic layers were dried over MgSO_4 and

concentrated. The residue was purified by flash chromatography on silica gel (petroleum ether / dichloromethane: 8/2).

E-1,2-bis(3-(diphenylphosphino)phenyl)diazene E-3. (66% yield, 84 mg, orange solid). ^1H NMR (500 MHz, CDCl_3) δ 7.86 (d, $J = 7.6$ Hz, 2H), 7.81 (d, $J = 7.6$ Hz, 2H), 7.46 (t, $J = 7.5$ Hz, 2H), 7.42-7.30 (m, 22H); ^{13}C NMR (75 MHz, CDCl_3) δ 152.7 (d, $J_{\text{C-P}} = 6.9$ Hz, C), 139.1 (d, $J_{\text{C-P}} = 13.2$ Hz, C), 136.8 (d, $J_{\text{C-P}} = 10.7$ Hz, C), 136.2 (d, $J_{\text{C-P}} = 18.2$ Hz, CH), 134.0 (d, $J_{\text{C-P}} = 19.7$ Hz, CH), 133.8 (d, $J_{\text{C-P}} = 20.0$ Hz, CH), 129.4 (d, $J_{\text{C-P}} = 6.9$ Hz, CH), 129.1 (CH), 128.8 (d, $J_{\text{C-P}} = 7.0$ Hz, CH), 122.2 (CH); ^{31}P NMR (202 MHz, CDCl_3) δ -4.9; HRMS (ESI) calcd. for $\text{C}_{36}\text{H}_{29}\text{N}_2\text{P}_2$ [M+H]⁺: 551.1806; found: 551.1781.

E-1,2-bis(4-(diphenylphosphino)phenyl)diazene E-6. (69% yield, 87 mg, orange solid). ^1H NMR (500 MHz, CDCl_3) δ 7.85 (dd, $J_{\text{H-H}} = 8.4$, $J_{\text{H-P}} = 1.3$ Hz, 4H), 7.42 (dd, $J_{\text{H-H}} = 8.4$ Hz, $J_{\text{H-P}} = 6.9$ Hz, 4H), 7.39 – 7.32 (m, 20H); ^{13}C NMR (75 MHz, CDCl_3) δ 152.8 (C), 141.7 (d, $J_{\text{C-P}} = 13.5$ Hz, C), 136.7 (d, $J_{\text{C-P}} = 13.7$ Hz, C), 134.4 (CH), 134.1 (d, $J_{\text{C-P}} = 19.7$ Hz, CH), 129.2 (CH), 128.8 (d, $J_{\text{C-P}} = 7.2$ Hz, CH), 122.9 (d, $J_{\text{C-P}} = 6.8$ Hz, CH); ^{31}P NMR (202 MHz, CDCl_3) δ -4.6; HRMS (ESI) calcd. for $\text{C}_{36}\text{H}_{29}\text{N}_2\text{P}_2$ [M+H]⁺: 551.1806; found: 551.1767.

General procedure for the synthesis of complexes **E-4** and **E-7**.

2,2'-Thiodiethanol (0.15 mmol, 15 μ L) was added to $\text{NaAuCl}_4(\text{H}_2\text{O})_2$ (0.076 mmol, 30 mg) in 3 mL of degassed water at 0°C. After 5 min, a solution of phosphines **E-3** or **E-6** (0.036 mmol, 20 mg) in 2 mL of degassed CHCl_3 was added and the mixture was stirred for 40 min at 0°C. The reaction mixture was extracted with CH_2Cl_2 , dried on MgSO_4 and concentrated *in vacuo*. The residue was taken up in a minimal amount of dichloromethane and pentane was added until precipitation. The supernatant was removed and the solid was washed twice with pentane and dried *in vacuo*.

Complex E-4. (93% yield, 34 mg, orange solid). ^1H NMR (500 MHz, CDCl_3) δ 8.06-8.00 (m, 4H), 7.66-7.55 (m, 16H), 7.54 – 7.48 (m, 8H); ^{13}C NMR (75 MHz, CDCl_3) δ 152.4 (C), 136.7 (d, $J = 13.4$ Hz, C), 134.3 (d, $J = 13.9$ Hz, CH), 134.2 (d, $J = 17.3$ Hz, C), 132.4 (d, $J = 2.5$ Hz, CH), 130.3 (d, $J = 12.2$ Hz, CH), 129.7 (d, $J = 10.5$ Hz, CH), 129.6 (d, $J = 12.0$ Hz, CH), 129.5 (d, $J = 15$ Hz, CH), 125.6 (CH); ^{31}P NMR (202 MHz, CDCl_3) δ 33.6; HRMS (ESI) calcd. for $\text{C}_{36}\text{H}_{28}\text{AuN}_2\text{P}_2$ [M-AuCl₂]⁺: 747.1393; found: 747.1424.

Complex E-7. (90% yield, 33 mg, orange solid). ^1H NMR (500 MHz, CDCl_3) δ 7.99 (dd, $J = 8.7$, 2.0 Hz, 4H), 7.68 (dd, $J_{\text{H-H}} = 8.6$ Hz, $J_{\text{H-P}} = 12.6$ Hz, 4H), 7.62 – 7.54 (m, 12H), 7.55 – 7.47 (m, 8H); ^{13}C NMR (75 MHz, CDCl_3) δ 154.2 (C), 135.0 (d, $J_{\text{C-P}} = 14.7$ Hz, CH), 134.2 (d, $J_{\text{C-P}} = 14.3$ Hz, CH), 132.3 (d, $J_{\text{C-P}} = 2.3$ Hz, CH), 129.4 (d, $J_{\text{C-P}} = 12.0$ Hz, CH), 123.5 (d, $J_{\text{C-P}} = 12.8$ Hz, CH); ^{31}P NMR (202 MHz, CDCl_3) δ 33.3; HRMS (ESI) calcd. for $\text{C}_{36}\text{H}_{28}\text{AuN}_2\text{P}_2$ [M-AuCl₂]⁺: 747.1393; found: 747.1404.

Absorption spectroscopy.

Absorption spectra and irradiation experiments were performed in spectroscopic grade acetonitrile (Carlo Erba) previously degassed by argon bubbling. UV-vis absorption spectra were recorded on a Cary 5000 spectrophotometer from Agilent Technologies. Photoisomerization reactions were induced by a Hg/Xe Lamp Hamamatsu LC8 Lightningcure, selecting the spectral line by means of an appropriate interference filter (Semrock FF01-315/15-25 + FF01-320/40-25, FF01-335/7-25, FF01-370/10-25, FF01-438/24-25). For thermal isomerization kinetics measurements, a Peltier temperature controller system was employed and the solution was thoroughly stirred over the whole experiment. The kinetic constants of the thermal isomerization processes were extracted by exponential fitting of the absorbance at a single wavelength, according to a first-order kinetic model.

Computational details.

Ground state geometry optimization by density functional theory (DFT) was performed with the Gaussian16 package at the B3LYP level using LanL2DZ and 6-311G(d,p) (for **3**, **4**, **6** and **7**) or 6-311++G (for **1** and **1'**) basis sets for gold and all other atoms, respectively.⁵⁹ The absence of imaginary frequencies was checked on all calculated structures to confirm they are true minima. Time dependent density functional theory (TD-DFT) was performed at the CAM-B3LYP level using LanL2DZ and 6-311G+(d,p) basis sets for gold and all other atoms, respectively, to compute the 30 lowest singlet states. In all calculations, the effect of solvent (MeCN) was taken into account using C-PCM model. Natural transition orbitals (NTO) analysis was performed on relevant transitions.⁵⁷ Molecular structures, orbitals and computed data were visualized and analysed with Avogadro and Lumo.^{60,61}

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