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Packing Polymorphism Affecting the Optoelectronic Properties of a π -Conjugated Organic Compound

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KEYWORDS Polymorphism, DFT, NCI, organic semiconductor

ABSTRACT [1]Benzothieno[3,2-*b*]benzothiophene (**BTBT**) derivatives are widely employed as hole transport materials in organic field-effect transistors. The electronic properties of these

materials depend critically on the crystal packing, which depends on its turn on the choice of the attached functional group. With symmetrically attached alcohol chains the structure may display different packing modes depending on the number of CH₂ groups in the alkyl chain. The di-pentanol **BTBT** derivative has two polymorphs **I** and **II**, crystallizing from different solvents, with distinct packing modes and consequently different electronic properties. Whereas the conformational changes are very small between **I** and **II** and the hydrogen bonding networks in the structures are identical, the adjacent **BTBT** cores are differently shifted and oriented with respect to each other. It is shown by Density Functional Theory that polymorph **I** having unfavorable electronic properties is slightly more stable than polymorph **II**. This is caused by the much more attractive cross stacking between **BTBT** cores and **C5OH** chains in **I** than in **II**. The stacking in **II** originates rather from electrostatic interactions between the **BTBT** cores. The differences and resemblances with the packing modes of the di-butanol and di-hexanol derivatives are discussed. Both polymorphic forms **I** and **II** display negative uniaxial thermal expansion, but in different directions with respect to the packing of the molecules.

Synopsis

The present study describes the packing polymorphism of a benzothienobenzothiophene derivative containing symmetric hydroxy aliphatic chains of which one form is semiconducting and the other not. Density Functional Theory and Non-Covalent Interaction calculations are used to discriminate the two forms. They are also structurally compared with very similar compounds having one methylene spacer more or less in the aliphatic chain.

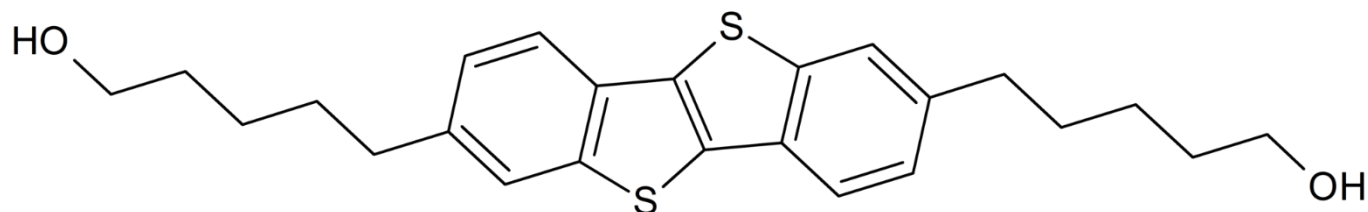
1. Introduction

Polymorphism is a widely occurring phenomenon in crystalline materials.¹⁻⁷ Physicochemical properties of polymorphs can be very different and thus having implications how they interact with their environment.^{8,9} Especially in pharmaceutical compounds the occurrence of polymorphism has been recognized as an inevitable process, which needs to be screened in an early stage of the development of new molecules in order to avoid undesirable effects when administering the drugs.^{10,11} The number of polymorphs is not limited to two, but can be as high as twelve as was recently reported for 5-methyl-2-[(2-nitrophenyl)amino]-3-thiophenecarbonitrile (also named ROY).¹²⁻¹⁵

A recent data mining study of the Cambridge Structural Database¹⁶ showed that at the time of writing 4573 distinct organic chemical compounds were flagged as ‘polymorphic’ in the CSD, of which 75% had more than 1 polymorphic structure described or reported. A remarkable conclusion from this study was that polymorphic entries within the CSD decrease, as a percentage, temporally, reflecting at the same time the huge number of new compounds being discovered and deposited at the CSD. This confirms in some sense McCrone’s thesis¹⁷ that the discovery of polymorphs is positively correlated with the energy and time put into the study of the compound, or in other words, most single crystalline compounds are not investigated in depth anymore whenever the growth of single crystals suitable for X-ray diffraction analysis appears to be successful. Trials with other solvents are not systematically carried out and growth conditions are not systematically varied, which makes the chance of finding polymorphs smaller. Only concomitant polymorphs may escape to this tendency, although the crystallographer may pick up only the ‘best’ crystals and ‘forget’ other tinier or differently shaped crystals.

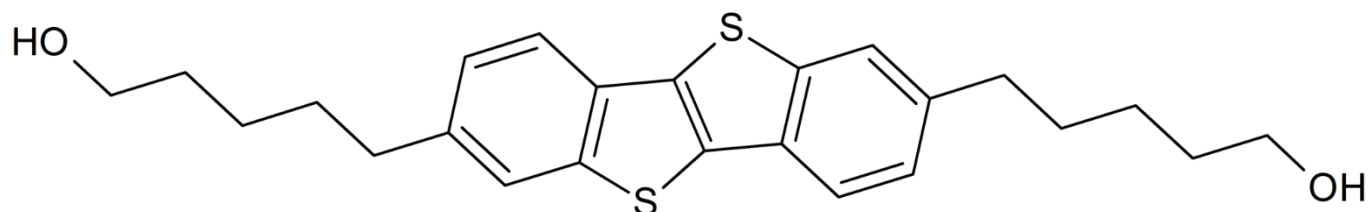
Another important finding in the Kersten study¹⁶ was the relative important propensity of multi-component systems prone to polymorphism. Steed & Steed¹⁸ and Clegg¹⁹ discussed the relationship between so-called $Z' > 1$ structures, *i.e.* structures with more than one independent molecule in the asymmetric part of the crystallographic unit cell, and the incidence of polymorphism.

Polymorphism is not only important in pharmaceuticals, but also in other fields. We focus here on organic π -conjugated compounds with excellent semi-conducting properties for organic field-effect transistor applications. We present a case of a benzothienobenzothiophene (**BTBT**) based derivative disubstituted by hydroxy aliphatic chains (pentanol) in positions 2 and 7 with two different polymorphs **I** and **II** (



Scheme 1), of which one is semiconducting and the other not. **BTBT-C5OH**, 5,5'-([1]benzothieno[3,2-*b*][1]benzothiophene-2,7-diyl)bis(pentan-1-ol), is intermediate to **BTBT-C4OH**²⁰ and **BTBT-C6OH**^{21, 22} of which we have reported the detailed structural evolution as a function of temperature. These studies showed the existence of isosymmetric phase transitions in both compounds, colossal negative thermal expansion (NTE) in **BTBT-C6OH** and small NTE in **BTBT-C4OH**. Using *ab initio* molecular dynamic studies it was shown that the room temperature structure of **BTBT-C6OH** is in fact a dynamic mixture of two different phases (but not polymorphs). The structure determinations for **BTBT-C5OH** were carried out in the framework of a large study on the **BTBT-C n OH** family of compounds²³. The structure of **II** recorded at room

temperature was reported in that study, but not described in detail. It was noticed as a side remark that the odd members of the series were much more difficult to crystallize than the even members, reflecting also different physical properties for the odd and even members. Based on the relative stacking of the **BTBT** cores in the structures of the semiconducting $n=6-10$ members a similar stacking was expected for the $n=5$ member. The crystallization in chloroform gave high-quality crystals, but not the expected herringbone packing of the **BTBT** cores allowing a semiconducting behavior. Therefore, more effort was put into the crystallization of **BTBT-C5OH** under different conditions and with other solvents. Eventually, with tetrahydrofuran (THF) as solvent, crystals were grown with the same thin plate habit as the $n=6-10$ members and X-ray diffraction analysis gave a similar crystal packing as found for the $n=6-10$ members. While this is a perfect illustration of McCrone's thesis,¹⁷ it also demonstrates that successful crystal growth is more than serendipity or luck and can be the result of mere perseverance.



Scheme 1. BTBT-C5OH

Polymorphism in the field of organic semiconductors can directly affect the desired application, *i.e.* an as high as possible charge carrier mobility, since the charge transfer integral that depends on the electronic wavefunction overlap is a highly sensitive function of the molecular packing.²⁴⁻
²⁸ This was demonstrated for another **BTBT** derivative, 2,7-dioctyl[1]benzothieno[3,2-*b*][1]benzothiophene, **BTBT-C8**, which, when deposited in the form of thin films using a novel

off-centre spin-coating method, displayed a higher mobility and a different molecular packing than the usual equilibrium structure.²⁹

In this study a detailed comparison is given between the structures of the two polymorphs of **BTBT-C5OH** and those of their $n=4$ and $n=6$ neighbors. Both low-temperature (LT) and close to room temperature (RT) structures are reported for both polymorphs. The Non-Covalent Interactions (NCI) descriptor (14) is used to analyze the packing differences between the polymorphs. Density Functional Theory (DFT) calculations are carried out to quantify the differences between the polymorphs and the charge carriers present.

2. Methods and materials

2.1 Crystal growth

The synthesis of **BTBT-C5OH** is reported elsewhere.³⁰ Polymorph **I** was obtained by slow evaporation of a saturated solution in THF disposed in a vial closed by a cap made of cotton at room temperature. Polymorph **II** was obtained during the cooling of a saturated solutions of **BTBT-C5OH** in chloroform (CHCl_3). The solution, in a closed vial, was heated before to be disposed in an oven at 60°C. Crystals could be obtained during the cooling to room temperature at a rate of 3°C per hour. Crystals of **I** appeared as colourless thick ($\sim 50 \mu\text{m}$, Fig. S1) millimetre sized platelets, whereas crystals of **II** grew also as plates, but much smaller and very thin ($\sim 1 \mu\text{m}$, Fig. S1).

2.2 X-ray diffraction

Diffraction data for the crystal structure of polymorph **I** were collected on a Rigaku-Oxford Diffraction Gemini diffractometer equipped with a Sapphire-2 CCD detector with sealed-tube Mo-

$K\alpha$ radiation at 115 and 275 K. The data collection was set up using *CRYSTALIS-PRO*³¹ and the integration of the data frames was done with the same program using default parameters. Lorentz and polarization effects were also corrected, and the empirical absorption correction was done using spherical harmonics employing symmetry-equivalent and redundant data.

Data for the crystal structure of polymorph **II** were collected on beamline ID29 of the European Synchrotron Radiation Facility (ESRF) in Grenoble, France, with an ARINAX-MAATEL MD2 diffractometer equipped with a Dectrix Pilatus-6MF detector. Data collection was performed with the *MXCUBE* program³² at 100 and 295 K and the raw image data were reduced with *XDS*^{33, 34} using default parameters.

The crystal structures were solved and the space group symmetry determined using the *ab initio* iterative charge flipping method with parameters described elsewhere³⁵ using the *SUPERFLIP* program.³⁶ Both structures crystallize in space group $P2_1/c$ and no change of space group was detected with temperature. One of the structures, **II**, has a β -angle close to 90° , and a check of possible twinning^{37, 38} indicated an 80% probability for pseudo-merohedral twinning and consequently an orthorhombic space group. This hypothesis was ruled out, because of a much higher R_{int} for orthorhombic symmetry and an unsatisfactorily refinement in the orthorhombic space group.

The structural models were refined against $|F|$ using full-matrix non-linear least-squares procedures as implemented in *CRYSTALS*³⁹ on all independent reflections with $I > 2\sigma(I)$. Table S1 gives full details about the data collections and the crystal structure refinements for the four data collections.

2.3 *Ab initio* computational methods

The starting point of the *ab initio* calculation of the two **BTBT-C5OH** polymorphs were the experimental LT unit cells containing 224 atoms and four **BTBT-C5OH** molecules. The

computation of the properties of very soft materials with a cohesive energy mostly dominated by van der Waals forces requires the use of a functional that includes dispersion contributions and the periodic electronic structure calculations were therefore performed using DFT within the PBE+D3 generalized gradient approximation including van der Waals contributions.^{40, 41} Projected augmented wave (PAW) pseudopotentials⁴² were utilized as implemented in the *VASP* code.^{43, 44} Γ -Centered ($1 \times 2 \times 2$) and ($3 \times 1 \times 2$) k -point meshes with a high k -point density in the direction of the smallest unit cell parameter were employed for the **I** and **II** polymorphs, respectively. The structures were computed using a 550 eV cut-off to account for the more electronegative oxygen and sulphur atoms and to properly describe the electronic polarization with the optimization of the positions of all or just a part of the atoms present in the unit cell, and with or without the optimization of the cell parameters. the residual forces on optimized atoms were lower than 0.01 eV \AA^{-1} for all calculations. The program *PLATON*^{37, 38} was used to recalculate the asymmetric unit content from the full unit cell content output resulting from the DFT calculations using the instruction CALC ADDSYM EXACT followed by ADDSYM SHX. Vibrations within the harmonic approximation were computed using finite displacements of 0.01 \AA to derive the dynamic matrix that was diagonalized to extract vibrational frequencies. Finite-temperature corrections and zero-point energies were calculated from these vibrations using a physical-statistical approach, and in order to obtain the temperature dependent free energy.

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3. Results

3.1 Crystal structure descriptions

Both polymorphs are $Z' = (0.5+0.5)$ structures, since they contain both two independent half-molecules in the asymmetric unit (**Figure 1**). Each independent molecule has internal inversion symmetry. A complete *MOGUL*⁴⁵ analysis with CSD 5.41 including the May2020 update⁴⁶ on bond lengths, valence and torsion angles, and ring geometry show two unusual features: the torsion angle C5-C6-C9-C10 in polymorph **I** at 115 K and 275 K with observed values 132.2(2) $^\circ$ and

134.5(4)° relative to an approximate mean database value of 90° for 4955 entries and the C6-C7-C8 valence angle for polymorph **II** at 295 K is slightly out of range as well with a value of 123.4(6)° compared to the mean value 121.07° with standard deviation 1.133° for 6520 database entries. Fig. S4a-b give the complete histograms resulting from the *MOGUL* analyses. The deviating torsion angle in **I** optimized by DFT at 0 K is 132.15° whereas the C6-C7-C8 angle in **II** is 122.05°. The deviating torsional angle can thus be considered as a genuine structural feature of **I**, whereas the deviating angle in **II** is probably due to a slightly less accurate structural model due to the very weak scattering power of this compound.

Interestingly, in both polymorphs the two independent molecules have a different orientation of the hydroxyl group with respect to the C₅H₁₀ alkyl chain: one is *cis* and the other is *trans*.

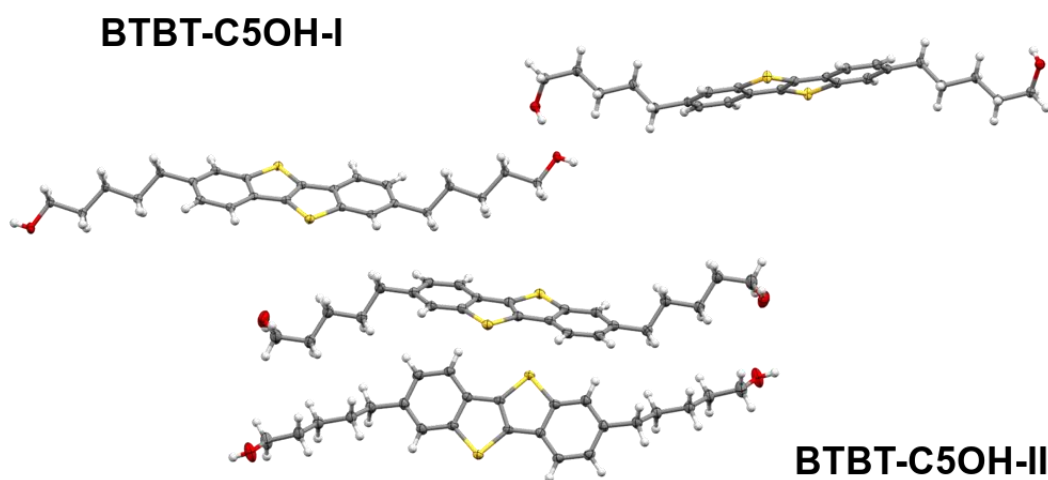


Figure 1. Atomic displacement (‘ORTEP’) plots of the polymorphs **I** (115 K) and **II** (100 K) of the structure of **BTBT-C5OH**. The upper two molecules represent the structure of **I** and the lower two molecules that of **II**. The atomic displacement ellipsoids are at the 75% probability level. Note the *cis* and *trans* conformations of the hydroxyl groups in the different molecules. All independent half-molecules have been completed by application of the internal inversion symmetry. Atom

labels are omitted for clarity; detailed atomic displacement plots with atoms labels of the two polymorphic forms are reported in Fig. S3.

The conformational flexibility of the independent molecules within each structure and between the two polymorphs can be assessed by calculating the best molecular overlap between the molecules. **Table 1** gives the RMS deviations of the overlapped molecules and maximum distances between the overlapped molecules calculated with *MERCURY*.⁴⁷ Between the two polymorphs there are four different overlap possibilities and between two temperatures within one polymorph two overlap possibilities.

Table 1. Molecular overlap RMS values and maximum distances between **I** and **II**

	I-II				
	rms	max			
LT	0.48	0.68	I	0.47	1.44
	0.69	1.06	II	0.57	1.84
	0.87	2.33			
	0.81	1.44			
RT	0.71	1.07	I	0.47	1.42
	0.80	1.42	II	0.55	1.78
	0.49	0.70			
	0.84	2.28			

Notes: All values are in Å. The values are given for the structures at low temperatures (LT) and close to room temperature (RT).

The results in **Table 1** show that the polymorphism is not of conformational nature and thus rather of the packing type. Within each phase the conformational changes with temperature are also rather small.

3.2 Crystal structure packings

Figure 2 shows the different packings of the structures of **I** and **II**.

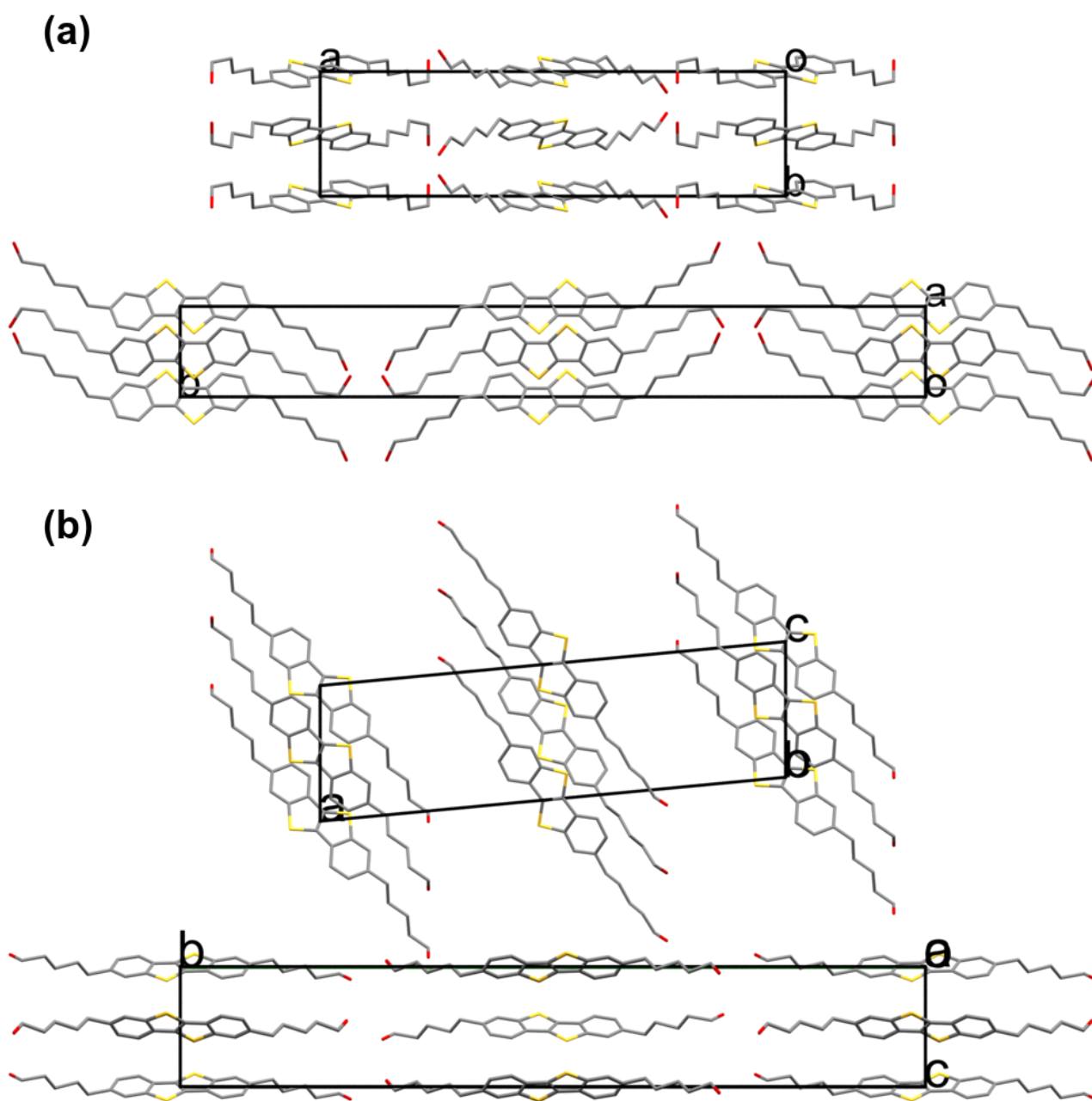


Figure 2. Crystal packings of structures **I** and **II** at low temperatures. The upper structure in each substructure is **I** and the lower **II**. The scale in each subfigure (a) and (b) is the same, but there is a slight scale difference between (a) and (b). The subfigure (b) was obtained from (a) by a 90° rotation about the horizontal x -axis of each polymorph.

The molecules of both polymorphs pack in layers perpendicular to the a^* axis for **I** and perpendicular to the b^* axis for **II**. The layers are connected via hydrogen bonds between the hydroxy-alkyl chains of different layers. The molecular arrangement within the layers is, however, completely different. In the structure of **II** the **BTBT** cores are spatially aligned and non-shifted in such a way that close intermolecular S \cdots S contacts exist of 3.619 Å at 100 K and 3.656 Å at 295 K which is only slightly larger than twice the van der Waals radius (1.80 Å, ⁴⁸) of sulphur. In the structure of **I** on the other hand, these close S \cdots S contacts are completely absent, since the **BTBT** cores of neighbouring molecules are laterally shifted along the axes of the **BTBT** cores. The interactions between the neighbouring molecules in **I** are mainly of the CH \cdots π type.

The orthogonal thermal expansion coefficients X1, X2, X3, and α_V calculated from the cell parameters at low temperatures and near ambient temperatures have been collected in **Table 2**. The thermal Indicatrix Anisotropy Coefficients (IACs; ^{20, 49}), which are a measure – between -1 and 1 – of the anisotropy of the thermal expansion, have also been calculated.

Table 2. Thermal expansion properties of the two polymorphs of **BTBT-C5OH**.

	X1	X2	X3	α_V	IAC	$\Delta\rho$
I	-38.34	52.90	152.27	167.13	0.38(-++)	-2.6
II	44.07	-48.56	125.29	120.27	0.52(+++)	-2.3

Notes: the values for the thermal expansion coefficients X1, X2, X3 and α_V are in MK⁻¹, where X1, X2, X3 are the three orthogonal principal axes of the thermal expansion tensor for the axes, and α_V the volumetric thermal expansion coefficients. X1, X2, and X3 are increasingly ordered by their absolute values; their sign is indicated within the parentheses following the IAC value. $\Delta\rho$ (%) is the decrease of the density between the lowest and highest measurement temperature.

It is interesting to note that both polymorphs display uniaxial negative thermal expansion of nearly equal magnitude: **I** within the *ac*-plane and **II** perpendicular to the *ac*-plane, which in terms of the molecular packing means that the negative thermal expansion in **I** is within the molecular planes, and in the case of **II** perpendicular to them. The cell parameters of **I** were also measured at intermediate temperatures and confirm the existence of the uniaxial negative thermal expansion (Fig. S2). The cell parameters of both structures were optimized at 0 K (see DFT section): the expansion along the axes which display negative uniaxial thermal expansion between 100 K and room temperature, show normal positive thermal expansion between 0 K and 100 K. This could indicate a crossover between a positive thermal expansion at low temperatures and negative thermal expansion at higher temperatures as is found for **BTBT-C4OH**.²⁰

Polymorph **II** is significantly less dense than polymorph **I**: 1.278 gcm⁻³ versus 1.314 gcm⁻³ near room temperature for **I** and **II**, respectively. This would imply, according to the well-known Burger & Ramberger ‘density rule’,⁵⁰ that polymorph **I** having the ‘best’ packing should have the lowest free energy at absolute zero, but it was shown in a recent study that this density rule is violated for almost half of the investigated monotropic cases,⁵¹ contrary to Burger & Ramberger’s 10% claim.

3.3 DFT structures

The first step in the DFT calculations performed at 0 K was to relax the positions hydrogen atoms while keeping the other atoms fixed as well as the cell parameters. The experimental X-ray structures of **I** and **II** at 115 K and 100 K, respectively, were used as starting points **O0** for the first optimisation **O1**. This led to C-H and O-H bond lengths very close to their neutron values of 1.089 and 0.993 Å, respectively, accompanied by a H-relaxation of about 300 meV per hydrogen and 36

eV per unit cell. Then all atoms were relaxed but the cell parameters were kept fixed (**O2**) and finally the cell parameters were also relaxed (**O3**). **Table 3** gathers the results of the optimizations **O0**, **O1**, **O2**, and **O3**, where **O0** is the starting point, *i.e.* the X-ray structures.

The computed energy changes from **O1** to **O2** for both polymorphs are lower than 200 meV per unit cell, *i.e.* less than 10 J.mol⁻¹/atom. This weak energy change already suggests that the calculated structures remains close to the experimental ones. The full unit cell optimization (**O2** to **O3**) leads to even weaker energy changes less than 30 meV/unit cell, *i.e.* 1.5 Jmol⁻¹/atom. These changes are in the expected range for phenomena linked to thermal expansion.

These calculations strongly suggest that polymorph **I** is slightly more stable than polymorph **II**, in agreement with the not so trustworthy Burger & Ramberger density rule.^{50, 51} Their relative energies are nevertheless very close with less than 1 meV per atom.

Table 3. Energies of the DFT optimizations **O0**...**O3**

	O0	O1	O2	O3	O3+ZPE	V(O3)	V(O0)
I	-	-	-1380.83	-1380.85	-1330.61	2047.57	2086.34
	1344.35	1380.64					
II	-	-	-1380.69	-1380.71	-1330.53	2112.97	2144.71
	1343.93	1380.58					
Δ_{I-II}	-0.42	-0.06	-0.14	-0.14	-0.08	-65.40	-58.37

Notes: See the main text for a definition of the different optimizations **O0**...**O3+ZPE (zero point energy)**. The energies are in eV. $V(\mathbf{O3})$ and $V(\mathbf{O0})$ are the cell volumes of the final optimizations and of the X-ray structures in Å³, respectively.

The DFT optimized structures were compared with the LT experimental structures via different structure similarity indicators, *viz.* the root mean square Cartesian displacement (RMSC,⁵²), the mean difference distance $\langle\delta\rangle$ between similar atoms and the measure of similarity Δ , defined by Bergerhoff *et al.*⁵³ and calculated with *COMPSTRU*⁵⁴ which performs the analysis in real space,

and the structure similarity parameter ϕ_{struc} defined in the *SUPERFLIP* program³⁶ which is rather a measure of similarity in reciprocal space since it measures the similarity between the structure factors calculated from the atomic structures.

Table 4. Similarity between DFT optimized structures (**O3**) and experimental low temperature structures.

	RMSCD1	RMSCD2	$\langle\delta\rangle$	Δ	ϕ_{struc}
O3_I(0 K)-I(115 K)	0.044	0.114	0.095	0.015	0.0050
O3_II(0 K)-II(100 K)	0.035	0.110	0.089	0.028	0.0036
I(115 K)-I(275 K)	0.051	0.093	0.078	0.053	0.0038
II(100 K)-II(295 K)	0.073	0.124	0.089	0.018	0.0169

Notes: See for a definition of the different similarity indices the main text. The similarity indicators between the low temperature structures and the room temperature structures for each polymorph are given in the two lower lines for comparison. Hydrogen atoms are included in the calculations in all cases, except for RMSCD1. The positions of the hydrogen atoms of the DFT structures have not been reset to their experimental X-ray positions.

Since the three parameters measure different similarities, it is not expected that they should give the same trend for the four different comparisons (**Table 4**). The purely geometric parameter $\langle\delta\rangle$ is fairly constant, whereas Δ , which takes into account as well the spontaneous lattice strain, shows more disperse values. Interestingly, the parameter ϕ_{struc} which measures the similarity in Fourier space, displays a rather deviating value for the diffraction patterns of the LT and RT phases of **II** without any obvious reason. Van de Streek & Neumann⁵² showed that RMSCD values for molecular single-crystal structures calculated without H atoms below 0.25 Å indicate in general ‘correct’ experimental structures. From the comparison with the RMSCD values calculated between the LT and RT experimental structures it may be deduced that the RMSCD values calculated between the 0 K DFT structures and the LT structures include a more or less important

temperature dependent component. The main conclusion from this comparison is that the experimental structures are without any doubt correct and that the DFT structures are a faithful representation of the experimental structures extrapolated to 0 K. Therefore, the DFT structures can be used to investigate the different **BTBT-C5OH**'s moieties contributions to the total crystal cohesive energy in the two different polymorphic structures.

3.4 Decomposition of the polymorph's cohesive energies

Because of the very small energy difference between the two polymorphs (less than 14 kJ/mol/unit cell), zero-point energies (ZPE) were computed from phonons and included in the energies of O3 (see Table 3). The energy difference between the two polymorphs decreases even more to only 7.7kJ/mol /unit cell (i.e. less than 2kJ/mol per molecule). The reduction in the energy difference is caused by polymorph **I** having slightly higher O-H frequency than **II** (e.g. the highest OH stretch frequency is 3287 cm⁻¹ for **I** while it is 3149 cm⁻¹ for **II**). This is consistent with the O-H distance of 1.00 Å in **I** and 1.01 Å in **II**. This also suggests that the intermolecular hydrogen bonds are slightly more strained and weaker in **I** than in **II**.

In order to investigate the different contributions to the total cohesive energy, the **BTBT-C5OH** molecule was optimized alone in a unit cell large enough to remove all the intermolecular interactions in order to obtain the cohesive energy (including the zero-point energy) of the free **BTBT-C5OH** molecule inside the structures of the two polymorphs. In this way the cohesive energy of the molecule (including zero-point energy effects) in the crystal can be computed to be -244.9 and -242.8 kJ.mol⁻¹ in the **I** and **II** polymorphs, respectively. Polymorph **I** appears to be slightly more stable than **II**, by about only 2 kJ.mol⁻¹ per molecule. The energy difference remains very low as can be expected for (kinetically) stable polymorphic structures.

In order to better understand the different energy contributions to the cohesive energy, we have partitioned the intermolecular interactions into three terms: the **BTBT/BTBT** intermolecular stacking interactions, the intermolecular **C5OH/C5OH** stacking interactions and the intermolecular cross-interactions between a lateral **C5OH** and **BTBT** cores. To obtain this decomposition, the **BTBT-C5OH** molecule is split into independent **BTBT** and **C5OH** subunits each of them saturated with hydrogens to compensate the broken C-C bond. Then, these subunits are separately computed in the frozen geometry of the **I** and **II** polymorphic unit cell with only the added saturating hydrogens being optimized. The crystals with only **BTBT/BTBT** and **C5OH/C5OH** subunits in the **I** and **II** structure are obtained and the associated interaction energy computed. Then, the different inter-moieties interaction energies are extracted as is shown schematically in **Figure 3**. In the two polymorphs, the **C5OH/C5OH** interaction is the main contribution ensuring about 50% of the cohesive energy as it includes the strong intermolecular H-bonding of the alcohol's moieties. Nevertheless, between **I** and **II**, the interaction contributions are quite different. In polymorph **II** the **BTBT/BTBT** cores stacking is 25 % stronger than in the structure of **BTBT-C5OH-I** and the **C5OH/C5OH** stacking 16 %. However, the cross stacking between **BTBT** cores and **C5OH** chains is three times more stable in **I** than in **II**. Therefore, in **I**, the softer coupling between adjacent **BTBT** cores and adjacent **C5OH** chains than that in **II** is more than fully compensated by the cross-stacking between **BTBT** and **C5OH** chains which gives an overall slightly more stable structure for **I** than for **II**.

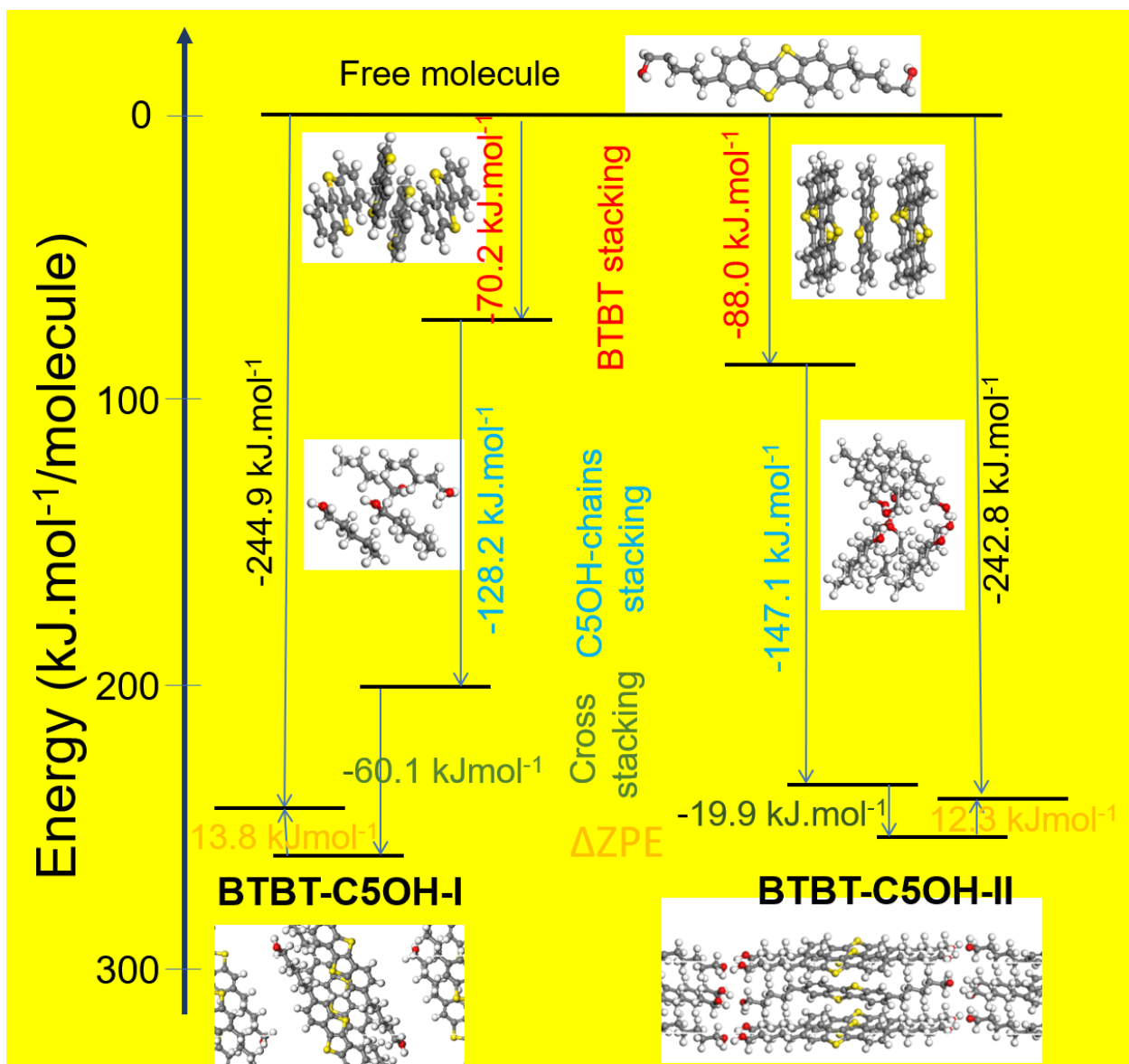


Figure 3. Decomposition of the different energy contributions to the total cohesive energy of the BTBT-C5OH molecule in the crystalline structures of **I** and **II**. The red values give an evaluation of the BTBT/BTBT intermolecular interaction energy, the blue ones the C5OH/C5OH intermolecular interaction, the green ones the cross-stacking energy between C5OH and BTBT moieties. The yellow values give the zero-point energies (ZPE) variation contribution to the cohesive energy. The black values give the crystal formation energy from BTBT-C5OH for the two different polymorphs **I** and **II**.

3.4 DFT computed electronic properties

The difference in molecular stacking in the **I** and **II** polymorphs is expected to modify the electronic properties of the two crystals. The electronic band structures of **I** and **II** were computed (Fig. 4) and are typical of direct semiconductors with a (DFT) band gap of about 2.2 eV. In each case, the lower conduction band and upper valence band are built from the Lowest Unoccupied Molecular Orbitals (LUMO) and Highest Occupied Molecular Orbitals (HOMO), respectively, of the four **BTBT-C5OH** molecules in the unit cell, without any band crossing. As seen in Fig. 4, most of the HOMO and LUMO orbitals of the **BTBT-C5OH** molecules are mostly developed on the π -system of the **BTBT** moieties with a small contribution of the orbital of the closest CH_2 groups that is associated with some classical electron donation/hyperconjugation behaviour of the aliphatic chains toward the **BTBT** moieties. The electron density variation from isolated subunits (Fig. S7) confirms that there is a charge transfer from the closest CH_2 groups in **C5OH** chains towards the **BTBT** cores associated to a donation/hyperconjugation effect. In the **BTBT** cores not all atoms are electron enriched: the 6 closer carbon atoms to the **C5OH** chains are enriched while other are even depleted (Fig. S7). The **BTBT-C5OH** molecule presents a strong electrostatic quadrupole with the CH_2 methylene group which is closest to the **BTBT** moieties being positively charged while the 6 aromatic carbon atoms that are within the **BTBT** core and closest to the chain are negatively charged. Therefore, dipole-dipole van der Waals interactions are expected to be quite important in the cohesive energy of these molecular crystals. Interestingly, in polymorph **I**, sliding of the relative packing compared to **II** allows a closer interaction between the negatively

charge atoms of the BTBT and the positively charged CH₂ groups leading to strong cross interactions that stabilizes **I** relative to **II** as found above.

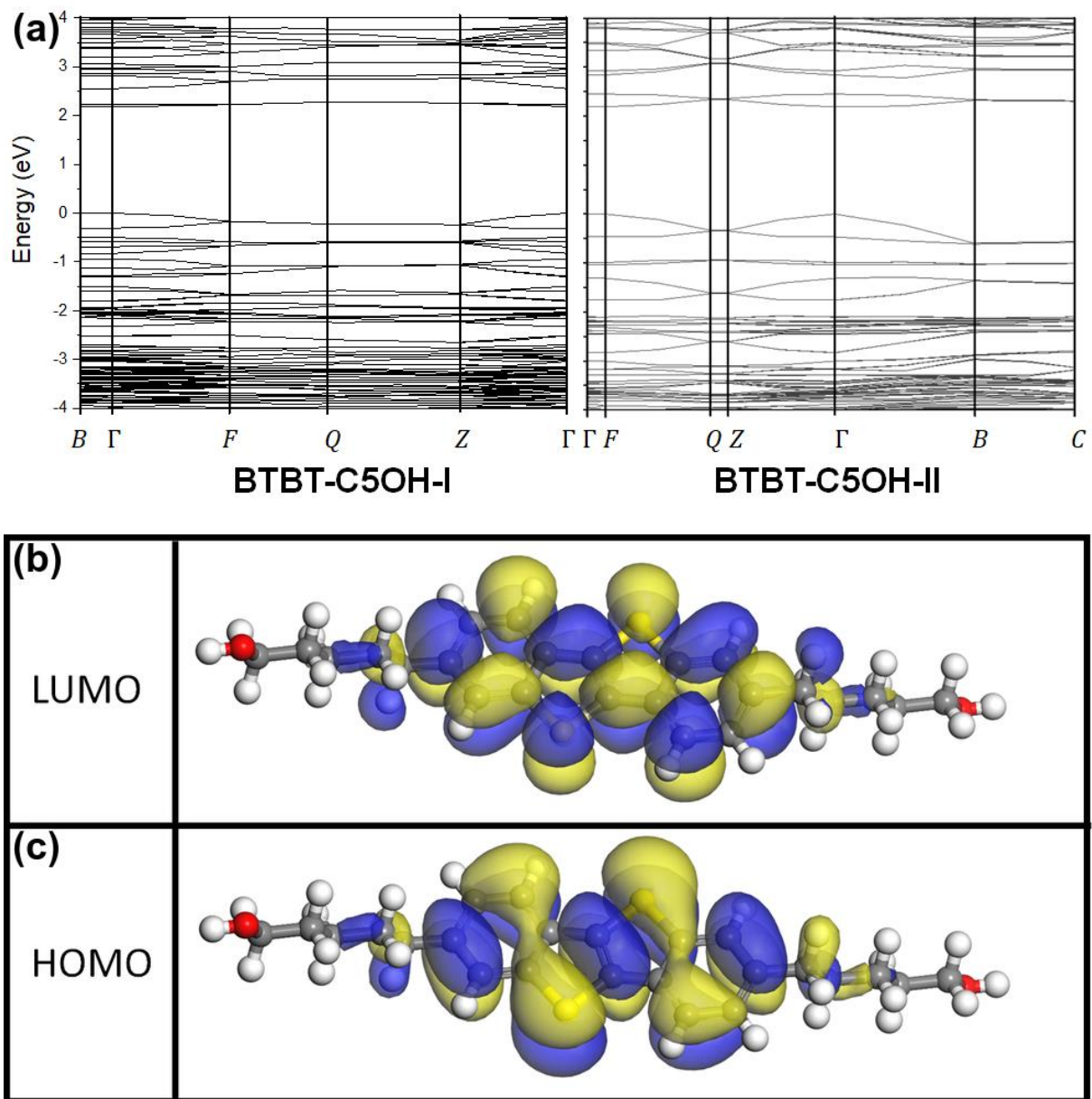


Figure 4. (a) Computed band structures of the **BTBT-C5OH-I** and **BTBT-C5OH-II** structures.

(b) LUMO and (c) HOMO orbitals of the **BTBT-C5OH-II** molecule.

The band structure clearly shows that some partial delocalization exists in the directions perpendicular to the **C5OH** chains mostly in the plane formed by the **BTBT** core. The band dispersion is more important for the polymorph **II** than for the **I**. This strongly suggests a far better molecular orbital overlap in the case of **II**, which is fully consistent with the structure of **II** where the **BTBT** cores are closely packed according to a herringbone mode. The Fukui functions f^- and f^+ that give insight in the real space delocalization of holes and electrons in the valence and conduction bands, respectively for polymorph **II** are presented in **Figure 5**. These functions show that electrons and holes are delocalized all over the **BTBT** core in the same molecular plane and they suggest that this polymorph is a decent organic semi-conductor with a good electron/hole mobility, as is indeed observed experimentally.³⁰

For **I**, as seen from the band structure (**Figure 4a**), the **BTBT** cores overlaps are not as strong because of the sliding of the molecule in the crystal, reducing the conjugation and thus electron/hole mobilities.

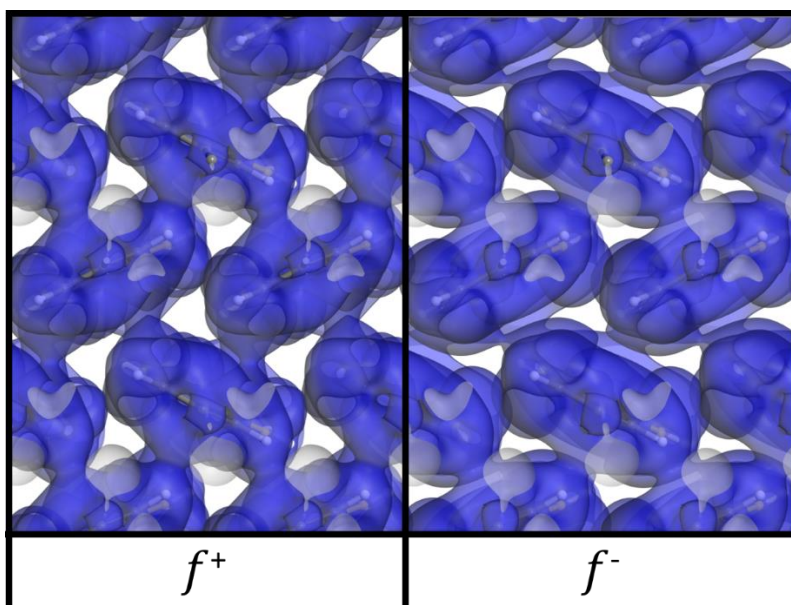


Figure 5. Positive contribution of the Fukui functions f^+ and f^- associated with the hole and the electron density in the conduction and valence bands respectively, for the **BTBT-C5OH-II** polymorph at the dibenzothiophene moieties.

3.5 NCI calculations

The Non-Covalent Interactions (NCI) method uses the scalar electron density function ρ derived from the atomic positions or from quantum-mechanical calculations to determine the reduced density gradient (RDG, $s(\rho)$) scalar field (12). Real space is in this way partitioned in different bonding regions depending on the values of ρ and s : low- s high- ρ indicate covalent bonds, low- s low- ρ to non-covalent interactions non-covalent interactions whereas high- s low- ρ correspond to non-interacting density tails. The sign of the second eigenvalue of the density Hessian matrix (λ_2) indicates whether the interactions are repulsive or attractive, where negative eigenvalues point to enhanced electron density close to the intermolecular bond critical points and positive eigenvalues represent steric clashes or non-bonding interactions.

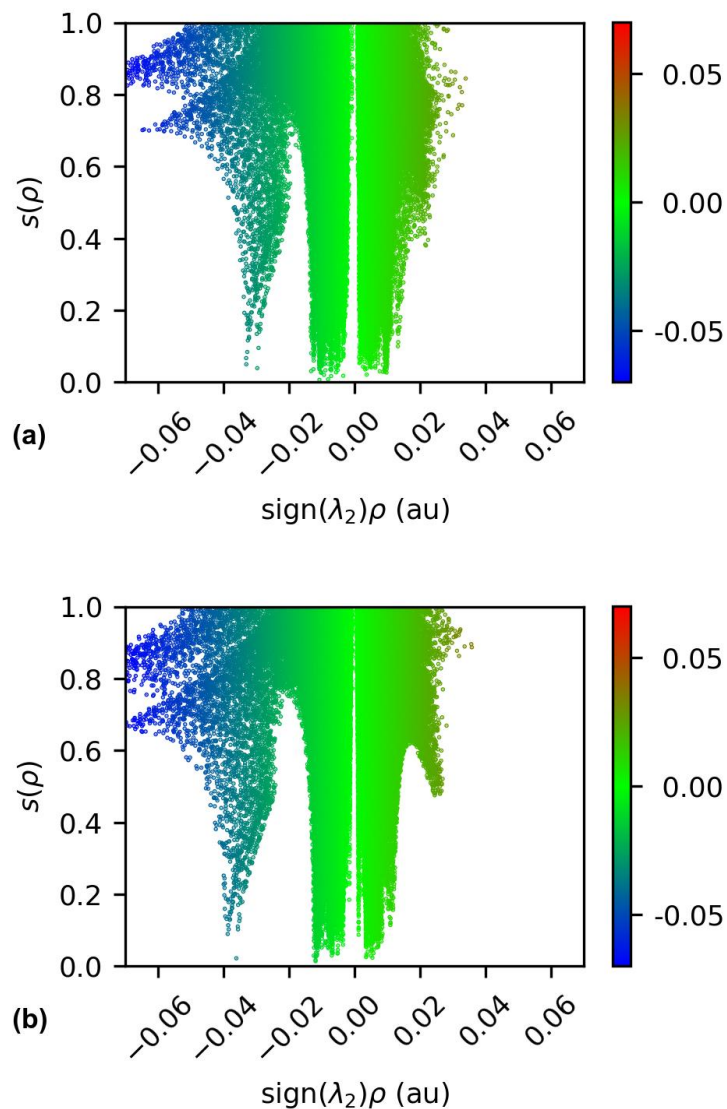


Figure 6. Reduced gradient (s) plots versus the promolecular density multiplied by the sign of the second eigenvalue of the density Hessian matrix ($\text{sign}(\lambda_2)\rho$) for polymorph **I** (115 K; a) and polymorph **II** (100 K; b). A simple colouring scheme based on the value of $\text{sign}(\lambda_2)\rho$ was taken, where blue represent attractive interactions (-0.07 au), green weak Van der Waals interactions (0.00 au), and red repulsive interactions (0.07 au).

Figure 6 gives $s(\rho)$ versus $\text{sign}(\lambda_2)\rho$ for the two polymorphs calculated from the promolecular densities, in which only intermolecular interactions were taken into account using the default

distinctive parameter $\gamma_{\text{ref}} = 0.95$. In both cases a cluster of six molecules (see Fig. S5) was taken for the calculations. Each cluster contains exactly 5 H \cdots OH interaction and four diffuse Van der Waals molecules. While the general features are rather similar, they do differ in the details. The blue hydrogen-bonding interactions appear to be slightly stronger in the case of polymorph **II** than in polymorph **I**, whereas the green weak Van der Waals interactions are also slightly more developed, especially on the positive side of $\text{sign}(\lambda_2)$.

The NCI regions can be quantified in terms of surface, area, and included charge. Boto *et al.* (12b) showed in particular that the integrals of the $n=2$, and $n=2.5$ powers of the charge over the NCI regions are highly positively correlated with the interaction energies determined from high-level DFT calculations. We determined here the NCI qualifiers for the different interaction regimes, i.e. the hydrogen bonding network for $\text{sign}(\lambda_2)\rho$ between -0.1 and -0.02 au, for the Van der Waals interactions between the alkyl chains and the BTBT moieties between -0.02 and 0.02 au, and for the repulsive interactions between 0.02 and 0.1 (nearly absent). Table 5 gives the relevant NCI properties for the two polymorphs within the different $\text{sign}(\lambda_2)\rho$ ranges. They show notably that both the volume of the NCI region and the integrated charge of **I** are significantly larger than for the less stable polymorph **II**. We also used the wave functions determined in the DFT calculations for a full periodic NCI calculation using the program *CRITIC-2* (Fig. S6),⁵⁵ but since they did not give additional insight compared to the promolecular cluster calculations, we do not detail them further.

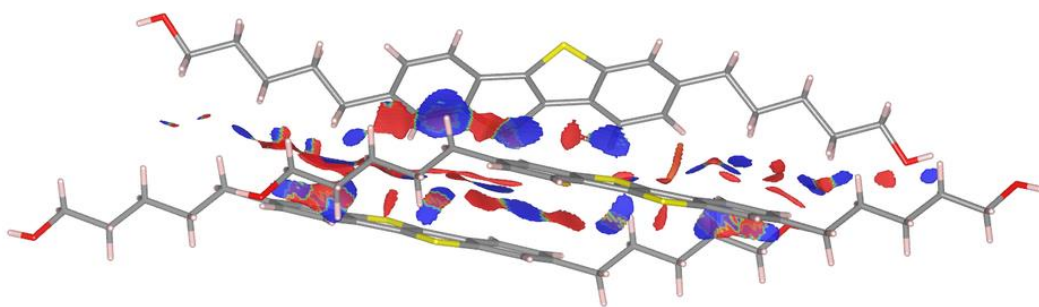
Table 5: NCI indicators for the structures of polymorphs **I** and **II**

	Polymorph I (115 K)		Polymorph II (100 K)	
$\text{sign}(\lambda_2)\rho$ (au)	I_0	I_1	I_0	I_1

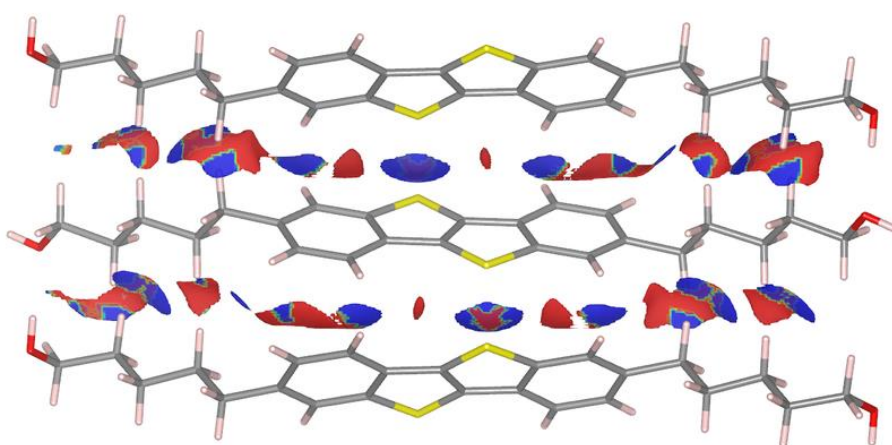
-0.1 -> -0.02	37.1	1.34	36.2	1.44
-0.02 -> 0.02	1464.0	10.0	1157.6	8.3
0.02 -> 0.1	2.9	0.08	1.9	0.05
all	1619.3	13.9	1288.8	11.9

Notes: I_0 is the volume Ω of the NCI region, I_1 is the integrated charge over the NCI volume Ω .

In order to determine spatially which Van der Waals interactions are the most attractive, the reduced gradient density isosurfaces were calculated for the Van der Waals interactions for $\text{sign}(\lambda_2)\rho$ values between -0.02 and 0.02 au (Fig. 7) at the 0.3 RDG isosurface level.



BTBT-C5OH-I



BTBT-C5OH-II

Figure 7. NCI volumes at the isosurface level $\rho = 0.3$ au. Upper figure: polymorph I (115 K); lower figure: polymorph II. The voxel distance used for the calculations is 0.075 au in all three directions. The following colour code based of $\text{sign}(\lambda_2)\rho$ was used: -0.02 au (blue) to 0.02 au (red). Note that this is not the same colour scheme as in **Figure 3**.

Figure 7 shows the NCI Van der Waals regions between trimers of **BTBT-C5OH** in the polymorphs **I** and **II**. The colour scheme goes from slightly attractive (blue) to slightly repulsive (red). The cohesive interaction in **I** seems to originate mainly from $\text{CH}\cdots\pi$ interactions between the aliphatic chains from one molecule and the aromatic **BTBT** core of an adjacent molecule,

whereas in **II** it is evenly composed from the contributions between thiophene moieties, $\pi\cdots\pi$ interactions between phenyls and interactions between aliphatic chains of adjacent molecules. This NCI interpretation is fully consistent with the energy decomposition obtained by DFT in the previous section.

3.6 Finite temperature effect

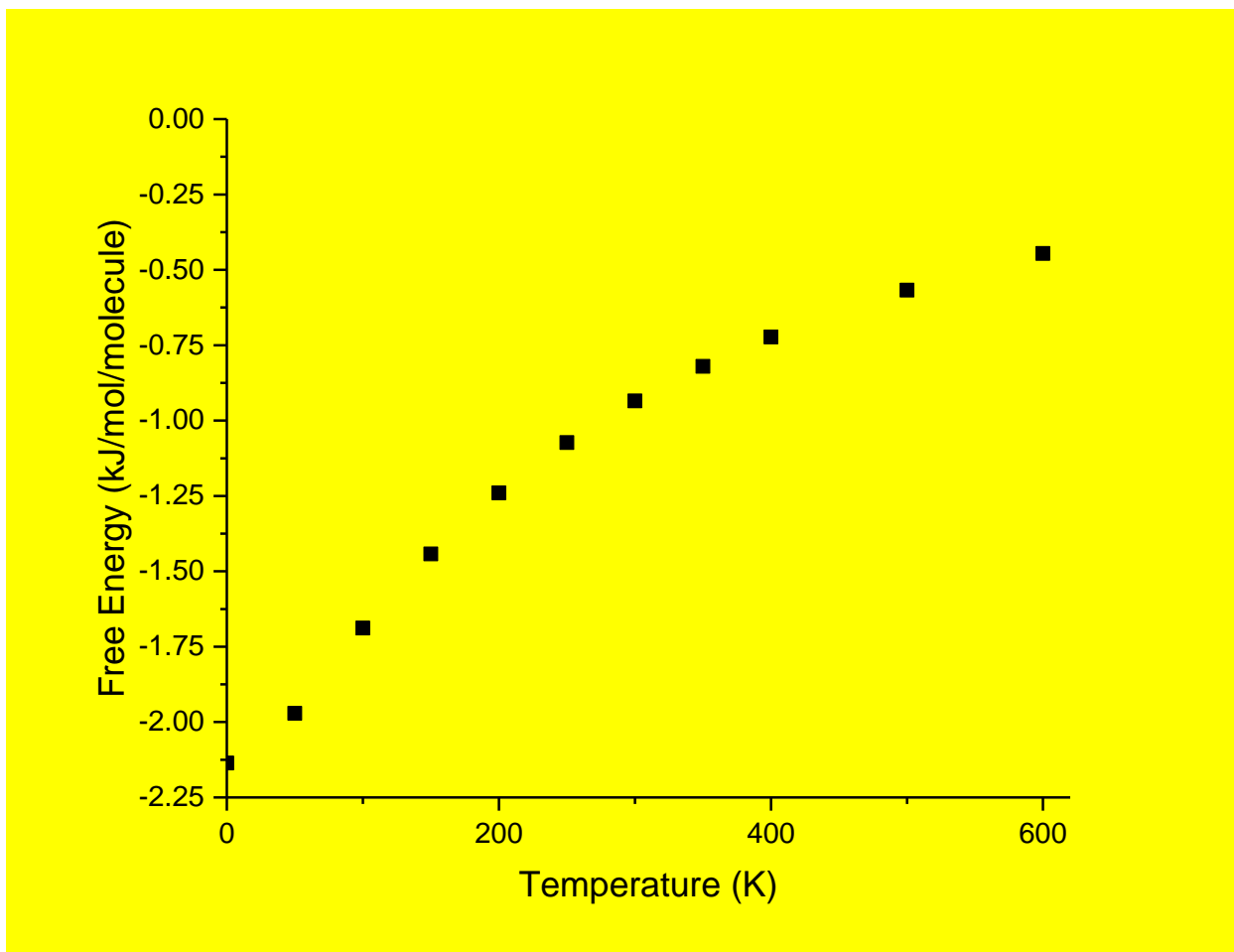


Fig. xxx Evolution of the Free Energy for the **II** ⇌ **I** polymorph conversion reaction using temperature dependant phonon enthalpic and entropic contributions within the harmonic approximation.

To investigate the finite temperature effect on the relative polymorphic stability associated with the $\text{II} \rightleftharpoons \text{I}$ equilibrium, we include the vibrational enthalpic and entropic contribution to compute the temperature dependant free energy within the harmonic approximation. The free energy is reduced by half from about -2 kJ.mol^{-1} to about -1 kJ.mol^{-1} between 0K and room temperature, but it hardly reaches -0.5 kJ.mol^{-1} at 600 K. The increase of the reaction free energy is the consequence of **II** having slightly more low energy vibrations that have larger entropic contributions at higher temperature than **I**. This could be probably linked with the slightly lower intermolecular interactions in polymorph **II**. Even if at high temperature, the approximation used to compute the free energy are not always accurate, it suggests that polymorph **I** is always the most stable and that stability inversion is not reached before crystal melting (in our model, stability inversion is found to be beyond 1000 K). Then polymorph **II** is purely a kinetic product that is not stable in the accessible temperature range. We also have used the computed free energy to estimate the solubility ratio of these two polymorphs as a function of temperature (see Fig. SXX in the SI). At room temperature the solubility ratio $S_{\text{I}}/S_{\text{II}}$ is calculated to be 0.42 meaning that **II** is more soluble than **I** as **II** is the less stable crystal form.

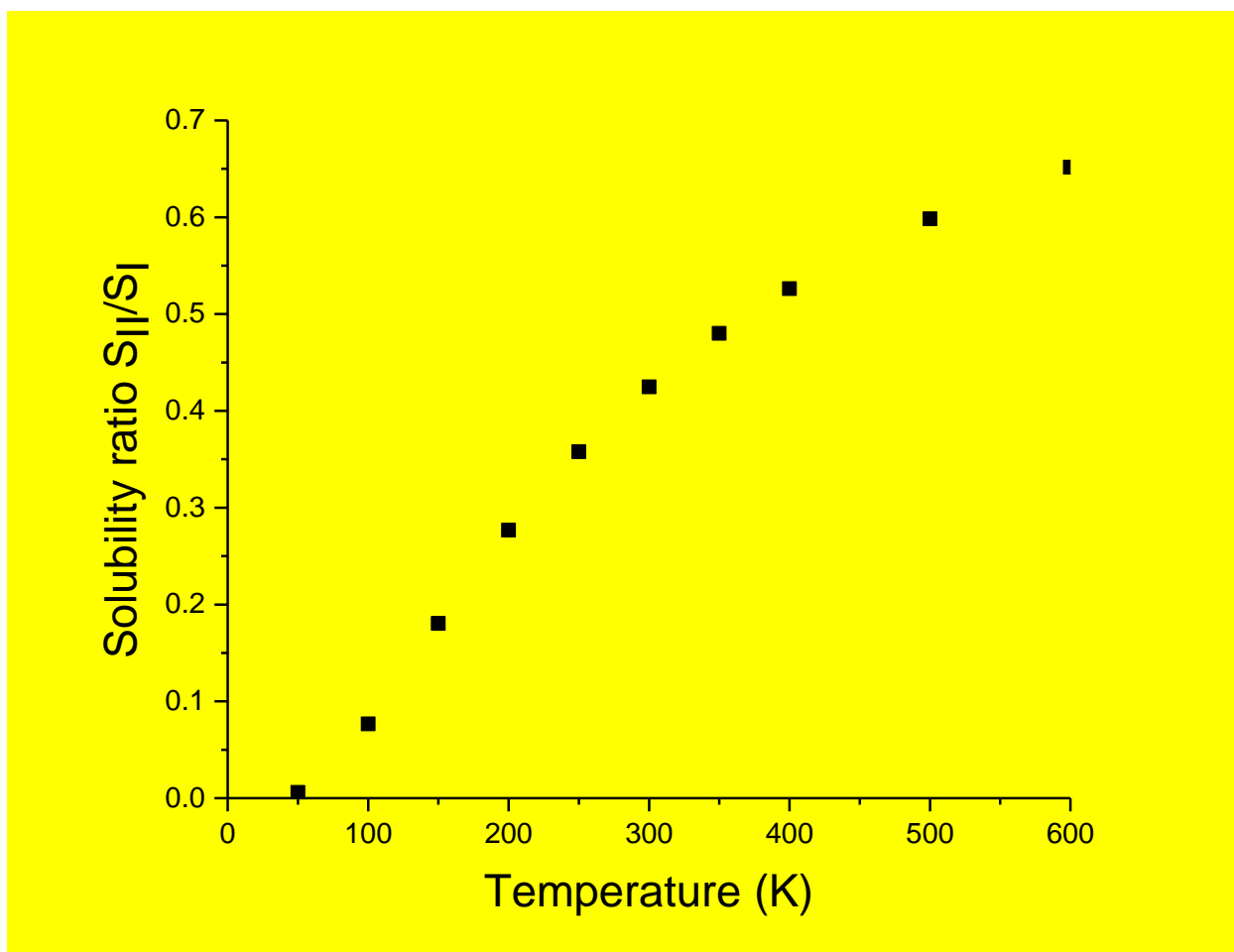


Fig Sxx: Evolution of the solubility ratio of polymorph II and I as a function of the temperature.

4. Discussion

A classification of polymorphs in four different types was proposed by Aitipamula et al. (1f), viz. conformational, synthon, packing and tautomeric polymorphism. As is clear from Table 1, the molecules within each polymorph and between the polymorphs are very similar, despite the torsional flexibility of the alkyl chains. The only important torsional flexibility which is observed between **I** and **II** concern the end alcohol group which is in *trans* configuration for **I** and in *cis* for **II**.

It is well recognized that in the case of synthon polyphormism arising from different hydrogen networks the less stable polymorphic form is often characterized by a less likely hydrogen bond

network than the more stable form has.⁵⁶ This type of polymorphism does not occur here, despite the different configurations of the alcohol groups, since the hydrogen bonding networks are very similar. This is corroborated by a hydrogen bond propensity analysis using a subset of CSD structures with acyclic aliphatic hydroxyl and cyclic thioether functional groups.⁵⁷ The H-bond propensity and H-bond co-ordination scores are - not surprisingly identical - for the two observed polymorphic groups (0.39/0.90) and more likely than any other possible network in the putative hydrogen bond network landscape, notably those involving the cyclic thioether functional group.

The *ab initio* DFT calculations suggest that polymorph **I** is slightly more stable than **II**. The fully optimized structure of **I** is about 150 meV lower in energy than that of **II**, which corresponds to less than 1 meV per atom. It is therefore not surprising that the two polymorphs exist. The NCI calculations and the DFT energy decompositions also suggest a slightly more stable structure for **I**, because of the larger accumulated charge between the herringbone stacked molecules than in **II**. The charge accumulated in the NCI regions in the hydrogen-bonded network is nearly the same for two polymorphs, so that the difference in stability can only be explained by the stacking mode of the molecules. The essential difference between the two polymorphic forms must thus be found in a difference between the weaker van der Waals interactions associated with the interactions between the slightly positively charged **C5OH** chains and the negatively charged **BTBT** cores.

Whereas the energy difference between the two polymorphs of **BTBT-C5OH** is only very small, the two individual forms have packing features that are more or less equivalent to other **BTBT** containing structures. The *MERCURY* packing similarity module was used with medium packing feature tolerances, i.e. 20° at maximum for intermolecular angles and 30% for intermolecular

distances. The **BTBT** cores in **I** are stacked in the same way as in **BTBT-C2** (RMSD = 0.323; 2,7-diethyl[1]benzothieno[3,2-*b*][1]benzothiophene,⁵⁸), and also to a lesser extent as the cores in 2,7-bis((5-methylhexyl)oxy)[1]benzothieno[3,2-*b*][1]benzothiophene (RMSD=1.718; ⁵⁹). The packing in **II** gives 14 hits with the same packing feature tolerances. Very close matches are found for **BTBT** cores symmetrically functionalized with C8 (RMSD=0.053; ⁶⁰), C10 (RMSD=0.044; ⁶⁰) and C12 (RMSD=0.03; ⁶¹) aliphatic chains. The ESI table S2 gives the complete listings of the matching structures with similar **BTBT** stacking patterns.

It is interesting to compare the **BTBT** stacking in the structure of **I** and **II** with that in the previously published symmetric structures of **BTBT-C4OH**²⁰ and **BTBT-C6OH**.²² Whilst there is no packing feature match between the **BTBT** cores in **BTBT-C5OH-I** and those in **BTBT-C4OH** and **BTBT-C6OH** there is a weak match between the **BTBT** stacking in **BTBT-C5OH-II** and that in **BTBT-C6OH**. Fig. 8 shows the best overlay between the **BTBT** cores of **BTBT-C5OH-II** (100 K) and **BTBT-C6OH** (95K; RMSD = 0.749 with intermolecular angle differences between 20° and 30° and intermolecular distance differences between 30 and 40%, thus establishing a weak packing similarity according to the *MERCURY* criteria). It is noted that this concerns only the stacking of the **BTBT** cores and not that of the complete molecules. The essential difference between the two packing features is the angle between the least-squares planes of the **BTBT** cores of the two structures which is acute (< 90°; 53.13°) for **BTBT-C6OH** and obtuse (between 90 and 180°; 124.77°) for **BTBT-C5OH-II** (**Figure 8**). This angle, also called the herringbone angle determines whether there are close S...S contacts between neighbouring **BTBT** moieties or not. Obtuse interplanar angles lead to closer S...S contacts and thus to better electronic properties than acute interplanar angles. An alternative but equivalent view is developed in Roche *et al.*³⁰ where it is shown that one half of the **BTBT** moieties for the even members in

the **BTBT-C n OH** series ($5 \leq n \leq 10$) with herringbone-type stackings is rotated by 180° with respect to the long axis of the moiety compared to the orientation in the structures of the odd members.

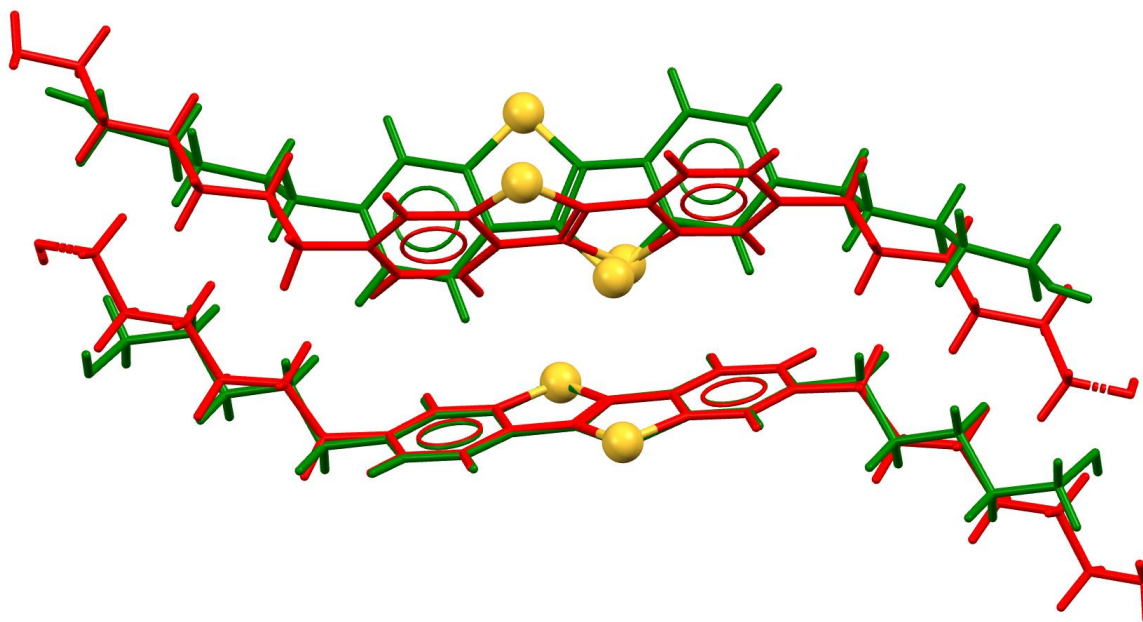


Figure 8. Packing similarity between the **BTBT** cores in **BTBT-C5OH-II** (100 K; green) and those in **BTBT-C6OH** (95 K; red). Both molecules are in stick representation except for the sulphur atoms. The **BTBT** cores of the lower pair of molecules have been aligned manually.

The average room-temperature mobility of 10 field effect transistor devices based on **BTBT-C5OH-II** fabricated on Si/SiO₂ substrates was reported to be $0.16(2) \text{ cm}^2\text{V}^{-1}\text{s}^{-1}$ (11), which does not make it to be an excellent semiconductor, compared to e.g. **BTBT-C8** ($1.1 \text{ cm}^2\text{V}^{-1}\text{s}^{-1}$, 40) or the organic semiconductor benchmark compound rubrene, C₄₂H₂₈, ($20 \text{ cm}^2\text{V}^{-1}\text{s}^{-1}$, ⁶²). It is known, however, that **BTBT** derivatives with longer side chains have in general better semiconducting properties (40).

5. Conclusions

Polymorphism remains to be a badly understood phenomenon. Crystal structure prediction has made large advances in the last decades,^{63, 64} but the bottleneck remains the prediction of almost equally stable polymorphs and how they can be realized.⁶⁵ The situation for **BTBT-C5OH** is not different: we are not able to say why the two polymorphs have very similar energies and why the one crystallizes in chloroform and the other in THF. A study as a function of pressure would be interesting to find out if the one form could be converted into the other.^{66, 67} This would at the same time induce an insulator to semiconductor transition.

BTBT-C5OH clearly presents a case of packing polymorphism, although there would have been the possibility of conformational polymorphism thanks to the rotational flexibility of the terminal hydroxyl group. However, in both polymorphic forms one of the two crystallographically independent molecules has two *cis* orientations of the hydroxyl group and two have *trans* orientations. The packing polymorphism in this case can be associated with slightly different distributions of weak intermolecular interactions between the two polymorphic forms as has also been demonstrated in some other cases.⁶⁸⁻⁷⁰

The present study shows once more the importance of polymorphism in the field of **BTBT** based organic semiconducting and gives in particular an insight into the role of the interactions between the aliphatic side chains and the cores, which, when neglected, can lead to polymorphic structures without sufficient overlap between **BTBT** cores and thus to decreased charge carrier mobilities.

ASSOCIATED CONTENT

Supporting Information. Crystal structure data of **BTBT-C5OH-I** and **BTBT-C5OH-II**, packing similarities between **BTBT-C5OH-I** and **BTBT-C5OH-II** and other structures in the Cambridge Structural Database, photomicrographs of crystals of polymorph **I** and **II**, thermal expansion data of **BTBT-C5-OH-I**, atomic displacement plots for the low-temperature structures of **I** and **II** with atom labelling, unusual geometric observations (Mogul analysis), 6-molecules clusters for polymorph **I** and **II** used for the promolecular NCI calculations, electron density variation between **BTBT-C5-OH** isolated subunits (DFT calculations). Crystallographic Information Framework (CIF) files for all structural models reported in this paper have been deposited at the Cambridge Structural Database under the numbers CCDC 2059375, 2060594-2060599. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/structures"

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Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Notes

The authors declare no competing financial interest.

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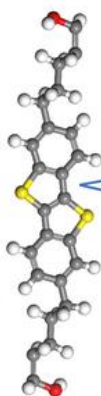
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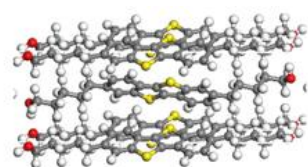
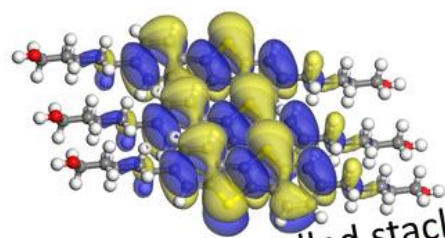
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Figure for Table of Contents

Investigating
polymorphic organization

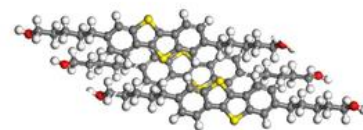
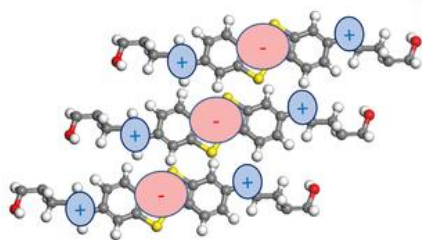


Orbital overlap controlled stacking



Strong electronic
delocalization

Electrostatic controlled stacking



Low electronic
delocalization