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# Influence of magnesium doping on microstructure, optical and photocatalytic activity of zinc oxide thin films synthesis by sol-gel route

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## **Abstract**

The article **deals with** the effect of Mg dopant on the properties and the photocatalytic of ZnO films synthesized by the sol-gel technique onto glass substrates. It is found that Mg dopant has a strong influence on the microstructure and functional properties. **The EDX analysis revealed that Mg had been incorporated into the ZnO.** Structural studies through XRD have shown that the deposited films are of hexagonal structure with a preferential growth along the c-axis **and nanometer crystal size**. According to AFM analysis; it is found that the roughness of Mg-doped ZnO films increases with the Mg dopant. The film surface depicted wrinkles morphology as observed by ESEM. Optical investigations by spectrophotometry and photoluminescence revealed that as the Mg dopant is increased, the optical band-gap increases. Moreover, near band edge and visible emissions as probed by Photoluminescence can be manipulated by the Mg dopant, as their intensity increases and decreases, respectively. Finally, the photocatalytic activity for degradation of methylene blue is strongly improved all the more as the Mg content is increased.

**Keywords:** ZnO; Sol-gel; Roughness; Microstructure; Band-gap; Photocatalytic test.

## 1. Introduction

Agro, paper and textile industries create effluents that are polluted with chemicals such as dyes, phenols, and pesticides. Some of these pollutants are recognized to be toxic and carcinogenic. Besides to that, they are not only harmful to the human but impact the ecosystem as well. Recently, society as well as regulatory authorities around the world, they became much interested in these chemicals due to their natural stability decomposition and their rising levels of by-products in the environment, mostly in water bodies. Unfortunately, the diversity of the available water treatment processes including adsorption, membrane filtration, and chemical treatment are insufficient to remove all the dyes. Conversely, these processes also create secondary contaminations like toxic gases and sludges that impose further treatment [1].

Moreover, the treatment of contaminated water with dyes and other organic compounds, photocatalysis based on nano-sized semiconductor photocatalysts is a recently emerging clean and cost-effective and elegant alternative for the treatment of water contaminated with dyes and other organic compounds [2]. The high photocatalytic performance of metal-oxide semiconductors, like titanium dioxide ( $\text{TiO}_2$ ) and zinc oxide ( $\text{ZnO}$ ), not only benefits from utilizing the high photosensitivity of these materials, but also their non-toxic nature, low cost, and eco-friendly nature [1]. As the photoactivity of catalysts as well influenced by the microstructural and chemical characteristics such as crystalline quality, size distribution, particle size, shape, phase composition and specific surface area [3], their optimization can be based on a materials science approach through microstructure and composition manipulation.

Zinc oxide ( $\text{ZnO}$ ) is an interesting catalyst material [4] as a semiconductor able to efficiently generate stable electron-hole pairs at room temperature. It an n-type semiconductor with large band-gap energy of 3.37eV and high excitation binding energy (60 MeV) making the exciton stable at room temperature [5]. Beside to that, it is also an attractive material as n-type semiconductor films in optoelectronic devices such as photodetectors [6, 7], solar cells [8], biosensors [9], gas sensors [10], and spintronic devices [11, 12]. The stable phase at room temperature is the well-known hexagonal wurtzite structure of  $\text{ZnO}$  [13].  $\text{ZnO}$  is also used in various other fields of applications such as optics, energy, chemistry, environmental applications [14, 15].

Furthermore, doping with metal or non-metal ions is one of the most effective approaches to promote the photocatalytic performance of ZnO. Metal doping can enhance the photocatalytic performance of ZnO by tuning the band-gap, introducing localized states, and modifying the morphology of ZnO. So far, a number of metal atoms, including, Al, Co, Mn, Ag, Fe, Mg, etc. [16] in the form of dopants were introduced into the ZnO crystal lattice to substitute Zn. An improvement in the doping technology of ZnO by Mg [17] has promoted many new applications in electronics and ultraviolet optoelectronic devices. Moreover, ZnO can be alloyed with high optical band-gap II–VI compounds like magnesium oxide (MgO) to raise its band-gap. The resulting new compound,  $\text{Zn}_{1-x}\text{Mg}_x\text{O}$ , has appeared as one of the most studied semiconductors due to its high band-gap. Rouchdi et al reported that many eminent researchers have analyzed the properties of Mg doped ZnO thin films, which showed that the structural, optical and electrical properties changed of ZnO after doping [6].

Thin films are of particular interest for photocatalysis as they are associated with a large surface/volume ratio, thereby limiting the possibility of electron-hole recombinations and the use of raw material. Therefore, good catalytic efficiency with a limited amount of matter can be achieved. In order to fabricate films, there are several methods such as sol-gel spin-/dip-coating [18, 19], spray pyrolysis [20], chemical bath deposition, pulsed laser deposition [21], and sputtering. Cost-effectiveness, thickness control, high transparency, versatility to synthesize multicomponent oxide layers of different compositions on various substrates, simplicity, excellent compositional control, homogeneity and lower crystallization temperature are special advantages for the fabrication of materials by the sol-gel process [22]. The sol-gel process allows the synthesis of materials of the same composition in various physical forms and coatings by changing experimental conditions.

In this paper, undoped ZnO and Mg-doped ZnO films were successfully prepared via a sol-gel procedure using a dip-coating process. The influence of magnesium addition on structural, morphological, optical properties and photocatalytic performance of ZnO films are investigated.

## **2. Materials and methods**

### **2.1 Synthesis of ZnO and Mg-doped ZnO**

Undoped and Mg-doped ZnO (Mg: 3, 5 and 7 wt. %) films were synthesized by a sol-gel method using dip-coating technique. The primitive solution was composed of precursor species: zinc acetate di-hydrate ( $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ ), a dopant source: magnesium acetate tetra-hydrate ( $\text{Mg}(\text{CH}_3\text{COO})_2 \cdot 4\text{H}_2\text{O}$ ), of a solvent: 2-methoxyethanol ( $\text{C}_3\text{H}_8\text{O}_2$ ) and of a complexing agent: monoethanolamine (MEA). First, zinc acetate di-hydrate 0.75 g were dissolved in 30 ml of 2-methoxyethanol with  $[\text{Mg}] / [\text{Zn}]$  weight ratios of 0, 3, 5 and 7 wt. %. The solution was heated under magnetic stirring at  $60^\circ\text{C}$  and 0.6 ml monoethanolamine (MEA) was added to accelerate the reaction. The obtained mixture was kept at  $60^\circ\text{C}$  for 2h under constant stirring. Then left free for 24 h to obtain a clear and homogenous solution. The substrates were cleaned in an ultrasonic bath using acetone, ethanol and distilled water, then dried at room temperature before the deposition of the films. Finally, films were deposited by dip-coating technique onto substrates of glass by the multiple dipping and drying at  $350^\circ\text{C}$  for 10 min. The coating to drying procedure was repeated ten times for each film. The multilayer films were finally annealed at  $500^\circ\text{C}$  for 1 h. **Figure. 1** show schematic of undoped and Mg-doped ZnO films synthesis using sol-gel method followed by dip-coating route.

The X-ray diffraction (XRD) data presented in this work were obtained using a PANalytical X'Pert Pro Philips diffractometer (using Cu  $K\alpha$  radiation source; wavelength 0.15406 nm). The morphology of the studied layers was identified by using Atomic Force Microscopy (AFM) using a nano-observer atomic force microscope developed by CS Instruments, and by Environmental Scanning Electronic Microscope type (ESEM) XL 30 FEG Philips. The transmittance was measured using a UV-3101 PC-Shimadzu double-beam spectrophotometer. The Photoluminescence (PL) experiments were performed using Fluorescence Steady-State Spectrofluorometer (Fluorologhoriba) with a 450W (Xe lamp as an excitation source).

## 2.2 Estimation of photocatalytic potential by the degradation of a dye

The photocatalytic activity of undoped and Mg-doped ZnO films was predestined by degradation of methylene blue dye (MB) under UV-light irradiation. An aqueous MB solution (50 mL) with an initial concentration of  $10^{-5}$  M was placed in a cylindrical glass vessel. Subsequently, undoped and Mg-doped ZnO films on the side were dipped into the petri dish, which contains the MB solution. Before the irradiation with UV-light source, the equilibrium of adsorption and desorption the

dark treatment of MB with photocatalyst was operated for 30 min, after that irradiated with light UV (365 nm 15W) at room temperature. The UV-light lamp was located at a distance of  $8.10^{-2}$  m above the top surface of the dye solution. The samples were taken out every 30 min from the reactor during irradiation, and the absorbance spectra were recorded using UV-3101 PC-Shimadzu double-beam spectrophotometer. The degradation efficiency of MB can be evaluated employing the following equation [23]

$$\text{Degradation efficiency} = \left( \frac{C_0 - C_t}{C_0} \right) \times 100\% \quad (1)$$

Where  $C_0$  is the initial concentration of dye solution,  $C_t$  is the concentration of dye solution after irradiation. The kinetics of the degradation has been demonstrated by Langmuir-Hinshelwood (L-H) model and can be written as [24]

$$r = \frac{dC}{dt} = \frac{KC}{1 + kKC} \quad (2)$$

Where  $K$  is the absorption coefficient of the MB,  $k$  is the rate constant,  $C$  is the concentration and  $r$  is the mineralization rate. When  $C_0$  is very small, L-H kinetics can be approached to first-order expression by the attached equations,  $k_{app}$  is the first-order rate constant.

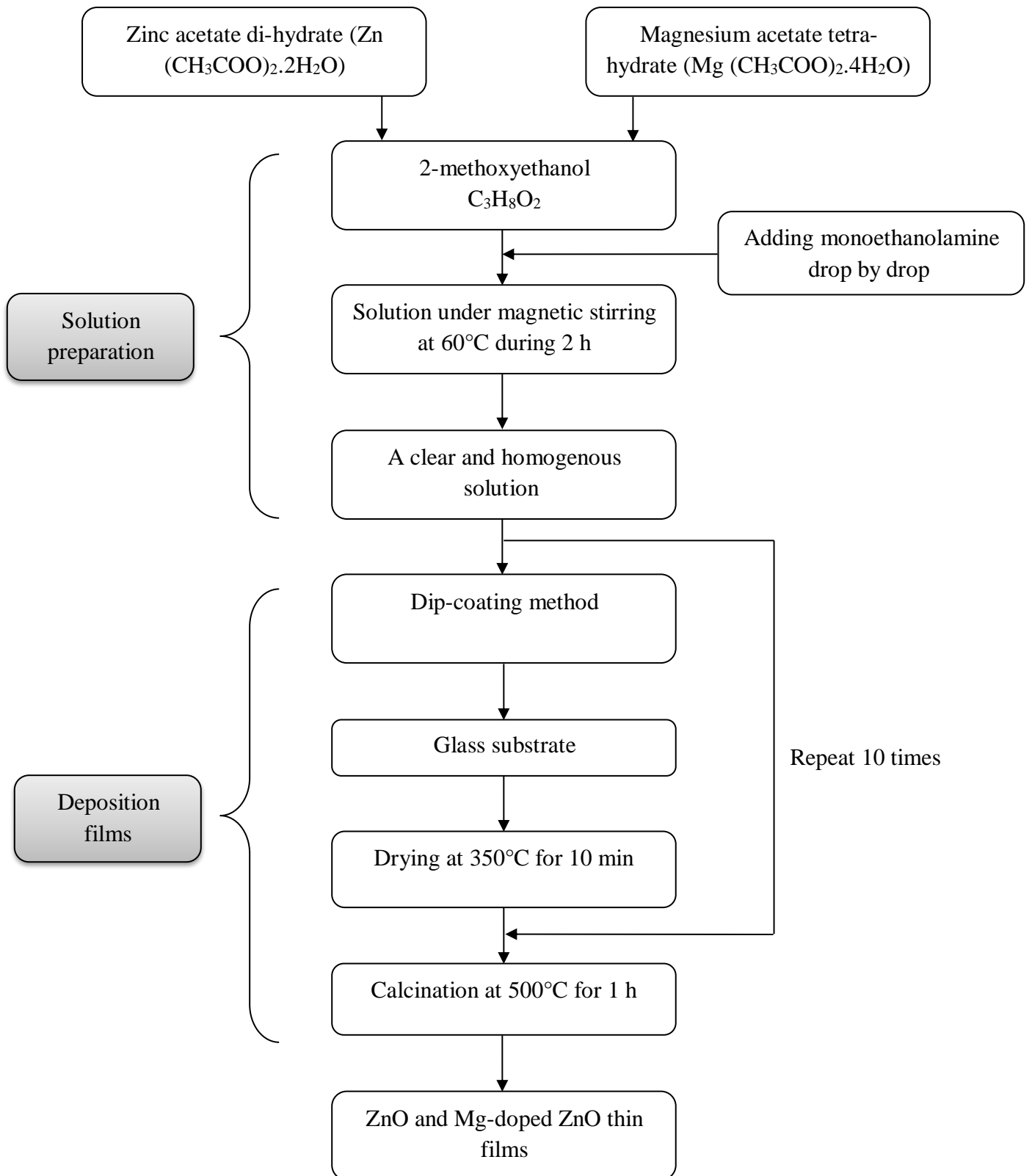
$$r = \frac{dC}{dt} = kKC = k_{app}C \quad (3)$$

Integrating the above equation,

$$C_t = C_0 e^{k_{app}t} \quad (4)$$

Or

$$\ln\left(\frac{C_0}{C_t}\right) = k_{app}t \quad (5)$$



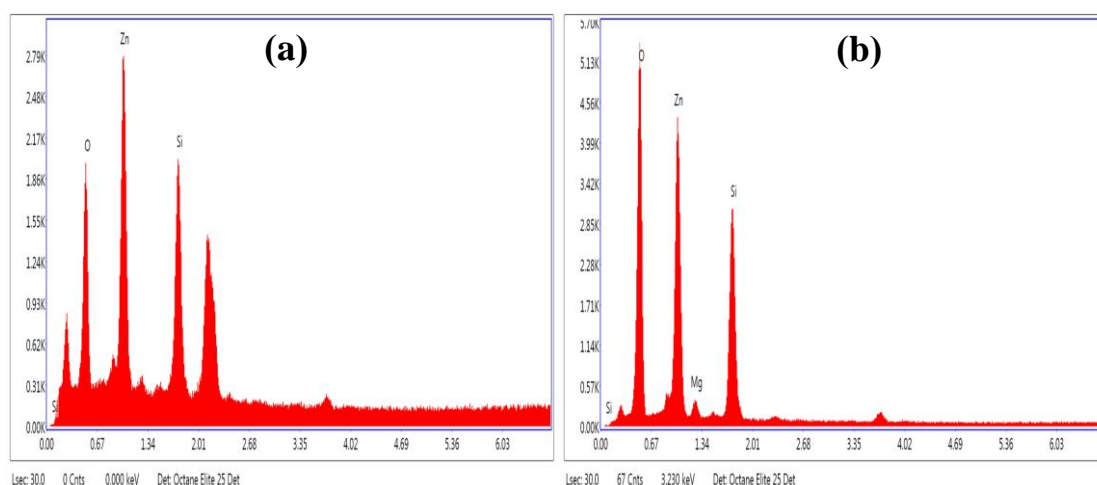
**Fig. 1** Synthesis diagram of undoped and Mg-doped ZnO thin films.

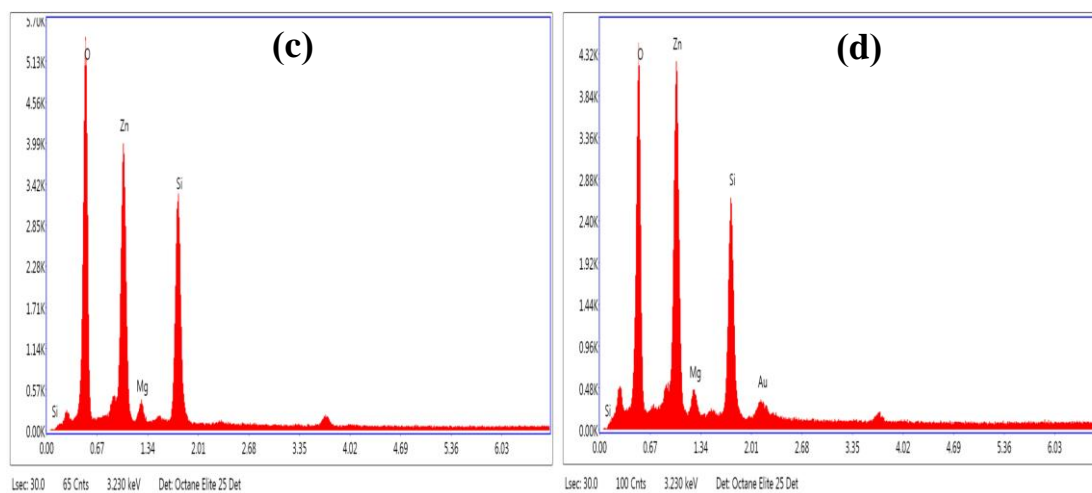


### 3. Results and Discussion

#### 3.1 Energy dispersive spectroscopy of X-ray (EDX)

The elemental analysis of undoped ZnO and Mg-doped ZnO thin films was analyzed by EDX. **Figure. 2** displays the EDX of ZnO and Mg-doped ZnO thin films. EDX spectra exhibit the peaks characteristic of the O-K, Zn-L, Mg-K transitions corresponding to film constituents as well as Au-M and Si-K transitions corresponding to glass substrate and thin gold deposition to avoid charge accumulation during measurement. The Semi-quantitative analyses from EDX measurement of the prepared films are presented in **Table 1**. The oxygen content is generally overestimated EDX. Moreover, some of the oxygen signal originates from the glass substrate. To estimate the Mg content in the film we have considered that it is half of the  $Mg/(Mg+Zn)$  ratio since the O/Mg and Zn/O ratios are supposed equal to 1. The estimated Mg content ranges between 0 for pure ZnO and 6.4 at. %. The measurements indicated that Mg is well incorporation into ZnO films using our sol-gel. Considering that only diffraction signal of ZnO was detected, we can consider that the process has well led to Mg doping. Even if the presence of secondary amorphous or very fine crystalline phase cannot be discarded, complete incorporation of Mg into ZnO is plausible knowing that the solubility limit of MgO into ZnO is in the range 5-8 mol % [25]. In the following the samples will be referred as ZnO, 4.9% Mg-doped ZnO, 6 % Mg-doped ZnO, 6.4 % Mg-doped ZnO, in order to consider the actual at. % Mg incorporated into the films.





**Fig. 2** EDX spectra of (a) undoped ZnO, (b) Mg-doped ZnO (4.9 at. %), (c) Mg-doped ZnO (6 at. %), and (d) Mg-doped ZnO (6.4 at. %) films.

**Table 1** Semi-quantitative analysis of EDX measurement of the prepared films

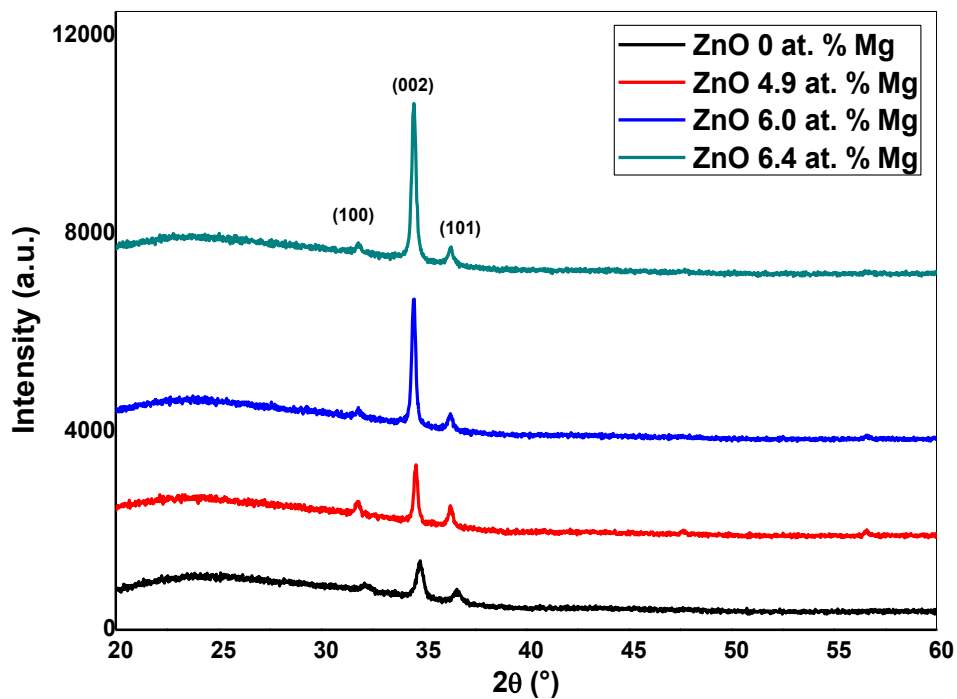
Composition of ZnO films	Element	Weight (%)	Atom (%)
Undoped ZnO	Zn	47.41	22.84
	O	21.47	42.25
	Si	31.13	34.91
Mg-doped ZnO 4.9 at. %	Zn	42.81	18.37
	O	32.14	56.35
	Mg	1.74	2.01
	Si	23.30	23.27
Mg-doped ZnO 6.0 at. %	Zn	40.12	16.83
	O	33.11	56.73
	Mg	2.03	2.29
	Si	24.74	24.15
Mg-doped ZnO 6.4 at. %	Zn	45.63	20.85
	O	28.74	53.65
	Mg	2.50	3.08
	Si	20.75	22.06
	Au	2.38	0.36

### 3.2 X-ray diffraction (XRD) analysis

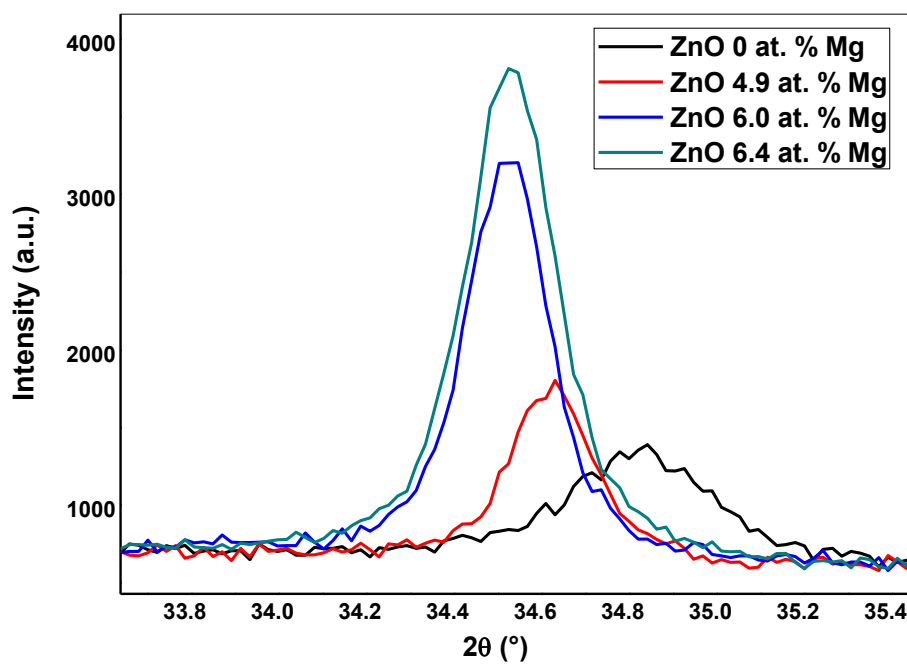
**Figure. 3** shows the XRD patterns of ZnO and Mg-doped ZnO films after the thermal treatment at 500°C for 1 h. All diffraction peaks (100), (002) and (101) were

indexed with the hexagonal wurtzite crystal structure of ZnO (JCPDS card no. 36-1451). From the XRD patterns, it is obvious that no peak of crystalline MgO or any other impurity phases could be detected. This suggests Mg is efficiently incorporated into the ZnO lattice without phase segregation occurs. The relative recorded diffracted intensities of the different peaks compared with the theoretical ones in a random configuration (JCPDS card no. 36-1451), indicate that the films are preferentially-oriented along the c-axis of the wurtzite structure as the signal for (002) peak dominates the diffractograms. This is all the more pronounced as the Mg is increases: the relative intensity ratio between (002) peak and other ones increases with the increase of Mg content. Whereas the width of the (002) intense peak decreases with the Mg concentration, indicating the average size of crystallites increases. **We note that the (101) plane has a higher intensity than (100). Rouchdi et al reported that the surface energy of (101) plane is the lowest that the surface energy of (100) plane of ZnO crystal [6].**

Additionally, the crystallinity of the ZnO phase **improved** with the addition of Mg into ZnO lattice as the overall diffracted intensity increases [26]. Notably, with increasing Mg dopant concentration, the position of the intense peaks (002) slightly shifted toward lower diffraction angles. **Figure. 4** shows the X-ray diffraction shifted in diffraction peaks (002) with Mg doping. This shift is due to the incorporation of Mg into the ZnO, **better highlights this shift. It corresponds to an expansion of the  $c(\text{\AA})$  lattice parameter and the same is observed for the  $a(\text{\AA})$  axis parameter by considering the shift in the position of the (100) peak [27]. Yet, the c-axis parameter of the undoped film is lower than that of bulk ZnO, suggesting pure ZnO film is in a tensile state before doping. Introduction of Mg into ZnO appears to relax this situation as the c-axis approaches to the bulk value upon Mg incorporation.**



**Fig. 3** XRD patterns of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films after the thermal treatment at 500°C for 1 h.



**Fig. 4** XRD shifted in the intense diffraction peaks (002) with Mg dopant.

The crystallite size  $D$  (nm) of undoped and Mg-doped ZnO films were determined using the Debye-Scherrer relation [28, 29]

$$D = \frac{K\lambda}{\beta \cos \theta} \quad (6)$$

Where  $K$  is the Scherrer constant value equals to 0.9,  $\lambda$  is the wavelength (0.154 nm),  $\beta$  is the full width at half-maximum (FWHM) intensity and  $\theta$  is the Bragg angle.

The residual in-plane stress  $\sigma$  (GPa) of films was calculated using the following relations [30]

$$\sigma = \left[ C_{13} - \frac{(C_{11} + C_{12})C_{33}}{2C_{13}} \right] e_{zz} \quad (7)$$

$$e_{zz} = \frac{c - c_0}{c_0} \quad (8)$$

Where  $c_0$  is the cell parameter of the bulk (taken at 0.520 nm),  $c$  is the experimental  $c$ -axis parameters,  $C_{ij}$  the elastic stiffness constants of ZnO ( $C_{11}=209.7$  GPa,  $C_{12}=121.1$  GPa,  $C_{13}=105.1$  GPa,  $C_{33}=210.9$  GPa).

The dislocation density  $\delta$  (lines/nm<sup>2</sup>) was calculated using Williamson-Smallman formula [31]

$$\delta = \frac{1}{D^2} \quad (9)$$

The results of the average crystallite size, dislocation density, the stress and the cell parameters were displayed in **Table 2**. It was noted that in case 4.9 at. % the average crystallite size increases whereas the dislocation density and stress decrease (see table 2), then when the Mg concentration increases up to 6 at. % there is a slight decrease in crystallite size and dislocation density while the trends of the stress continue to decrease, this behavior due to nucleation, growth and activation energy phenomena of the thin films and the substitute of  $Zn^{+2}$  by  $Mg^{+2}$ . Furthermore, in the case of 6.9 at. % the crystallite size reduces significantly whereas the dislocation and the stress increases, this may be attributed to the trend toward the formation of the secondary phase of MgO, due to  $Mg^{2+}$  which could leave the ZnO lattice, and work to return to a tensile state before doping addition.

**Table 2** Structural parameters obtained from XRD analysis and High score plus (data were taken from the intense (002) peak)

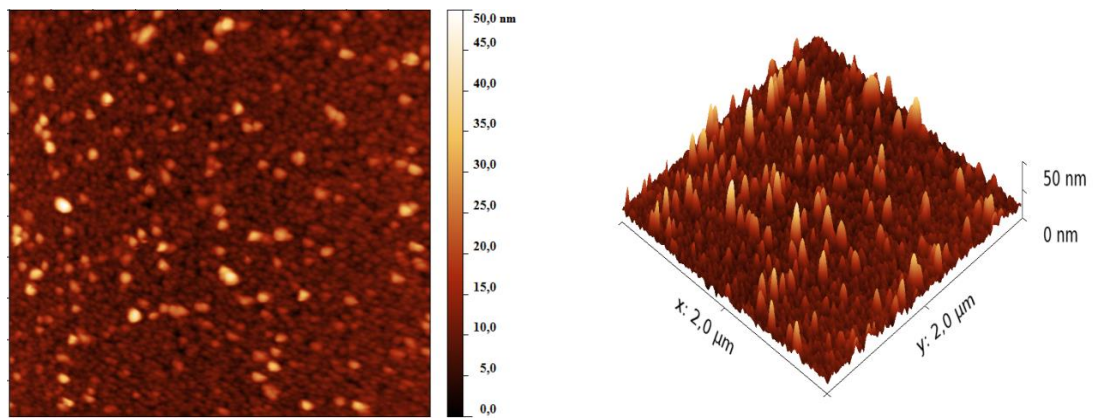
Samples	$\theta$ (°)	$\beta$ (°)	D (nm)	$\delta * 10^{-4}$ (lines / nm <sup>2</sup> )	$\sigma$ (GPa)	a(Å)	c (Å)
Undoped ZnO	34.79	0.36	24	17.36	2.17	3.21	5.15
4.9 at. % Mg-doped ZnO	34.49	0.18	51	3.84	0.74	3.25	5.18
6 at. % Mg-doped ZnO	34.49	0.19	50	4.00	0.04	3.24	5.20
6.4 at. % Mg-doped ZnO	34.50	0.22	42	5.66	0.08	3.25	5.20

### 3.3 Atomic force microscopy (AFM) analysis

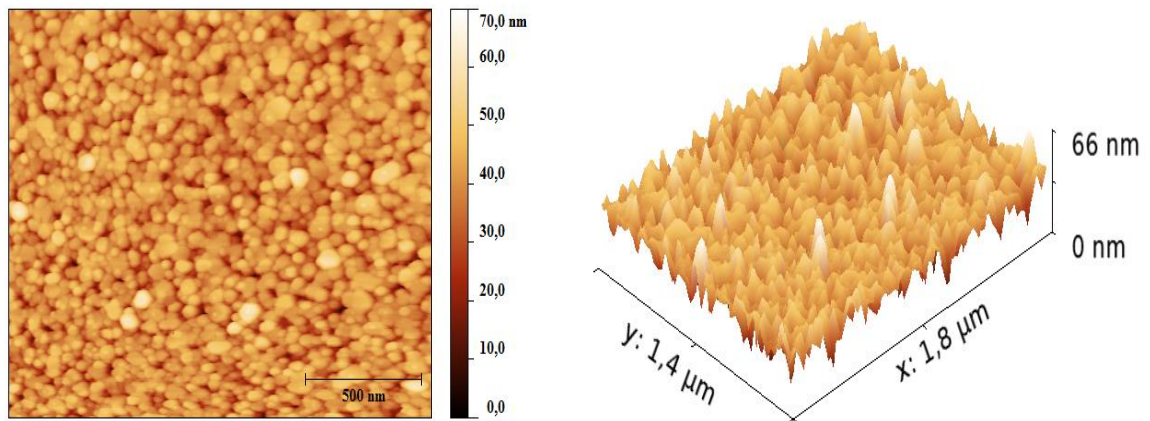
AFM is a useful technique to reveal the quality of the surface morphology and to estimate the lateral grain size and surface roughness. The two (2D) and three (3D) dimensional AFM images of undoped and Mg-doped ZnO films are shown in **Figure 5**. The root mean square roughness (RMS) **was** deduced from these images and values are given in **Table 3**. All films present a homogeneous structure with rounded shape grains, an increase and growth in grain width with Mg concentration. The RMS is found to increase with the Mg content; it changes from 4.46 nm to 12.30 nm for ZnO and 6.4 at. % Mg-doped ZnO, respectively, **means** that Mg-doped ZnO layers have a rougher surface than pure ZnO. This behavior can be related to ease grain growth upon Mg addition. This higher surface roughness and related larger developed surface area may be an advantage for photocatalysis, dye-sensitized solar cells (DSSCs), and gas sensing. **The thickness (d) was measured by using a profilometer instrument; the thickness (d) is** shrinking while raising the Mg content up to 6 at. % and then increasing slightly for 6.4 at. % Mg as represented in **Table 3**.

**Additionally, AFM images also can be used to evaluate the lateral grain size of the prepared samples, the films composed of nano-sized grains with wider grains for Mg-doped ZnO than undoped ZnO; the Mg-doped ZnO grain sizes were taken values stretch between 40-50 nm whereas undoped ZnO stretch 20 nm. ImageJ software was**

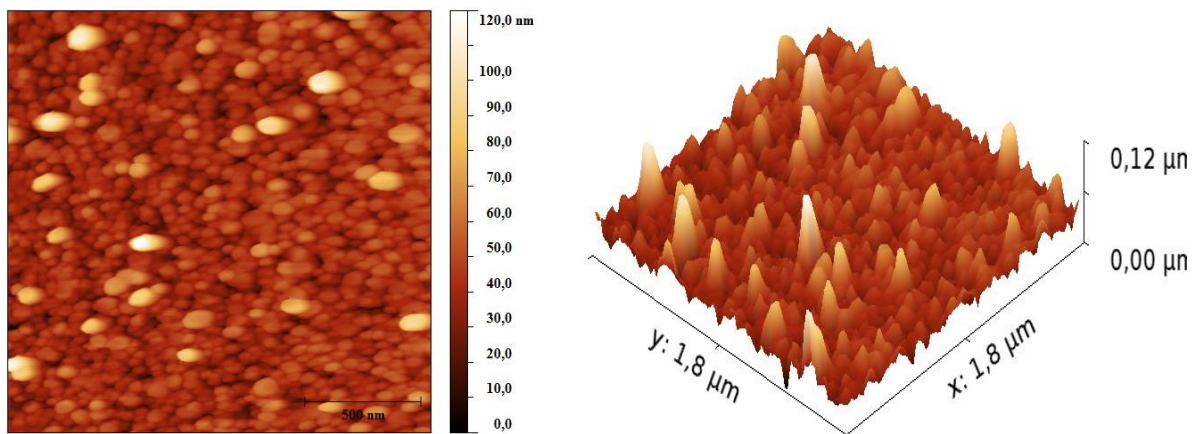
used to acquire the estimated values of grain sizes, these results are in agreement with those gained from XRD.



(a) ZnO 0 at. % Mg

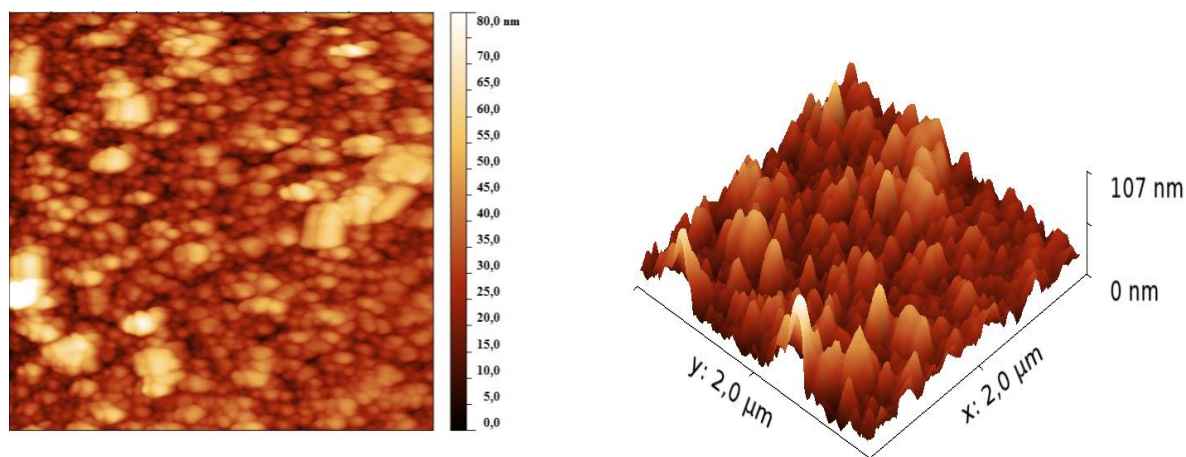


(b) ZnO 4.9 at. % Mg



(c) ZnO 6 at. % Mg





(d) ZnO 6.4 at. % Mg

**Fig. 5** 2D and 3D images of AFM of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films.

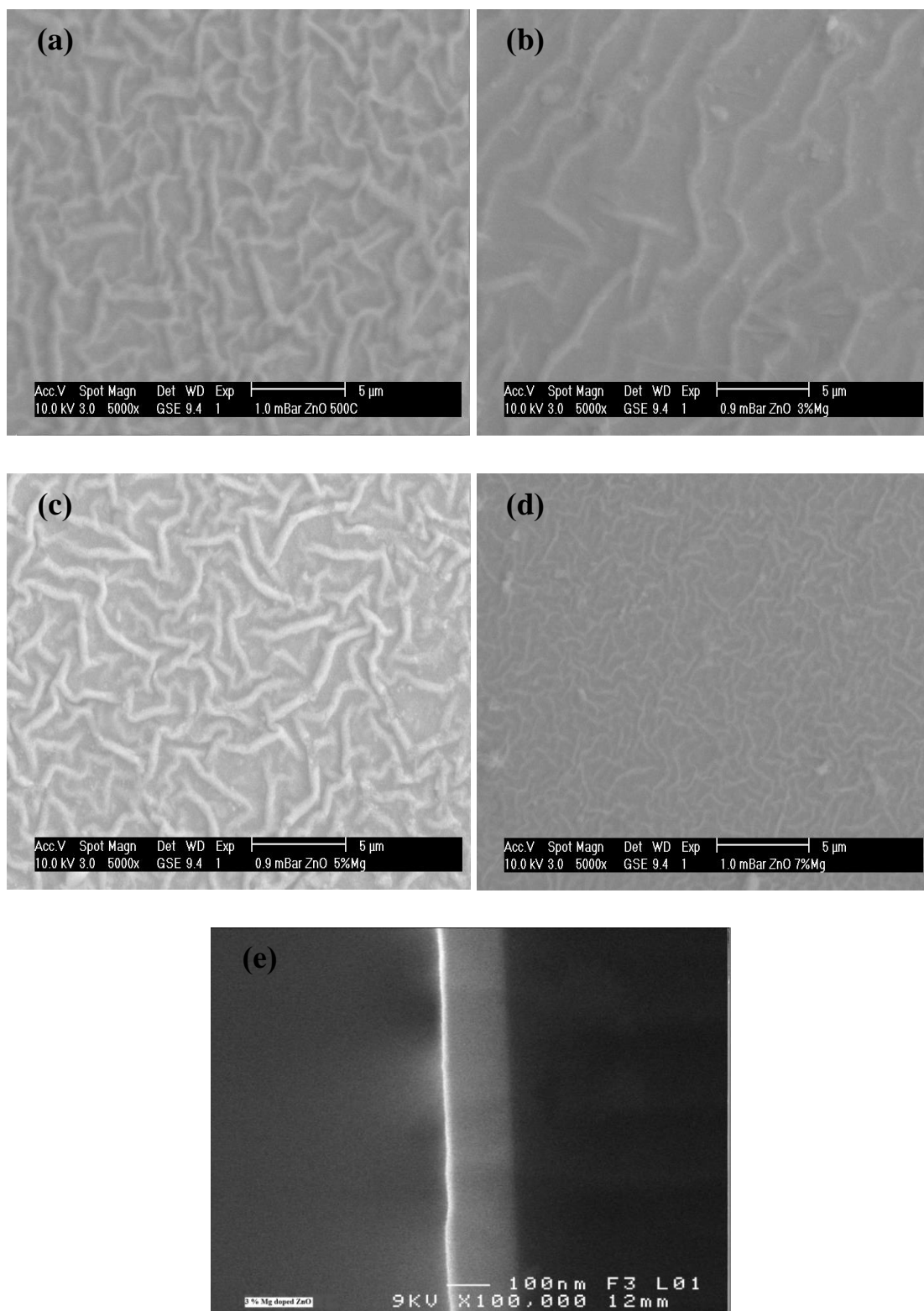
**Table 3** Variation of thickness (d), root mean square roughness (RMS) and Ra with Mg content.

Samples	Undoped ZnO	Mg-doped ZnO 4.9 at. %	Mg- doped ZnO 6 at. %	Mg-doped ZnO 6.4 at. %
d (nm)	331	161	134	199
RMS (nm)	4.46	6.48	12.20	12.30
Ra (nm)	2.87	5.00	8.40	9.40

### 3.4 Environmental Scanning Electron Microscopy (ESEM)

The morphology of the films was studied by the ESEM. The images of undoped and Mg-doped ZnO films and **cross section** obtained by sol-gel route are illustrated in **Figure. 6**. **These images indicate that the films are exempt of cracks or voids.** The surface of all the layers exhibits wrinkles and very small crystallites; these wrinkles are linked with each other and with the addition of magnesium to the layers, the wrinkles are shrinking and their density increases. **The presence of these wrinkles developed the surface area of the film larger than the projected surface area, which should be beneficial for catalysis.**

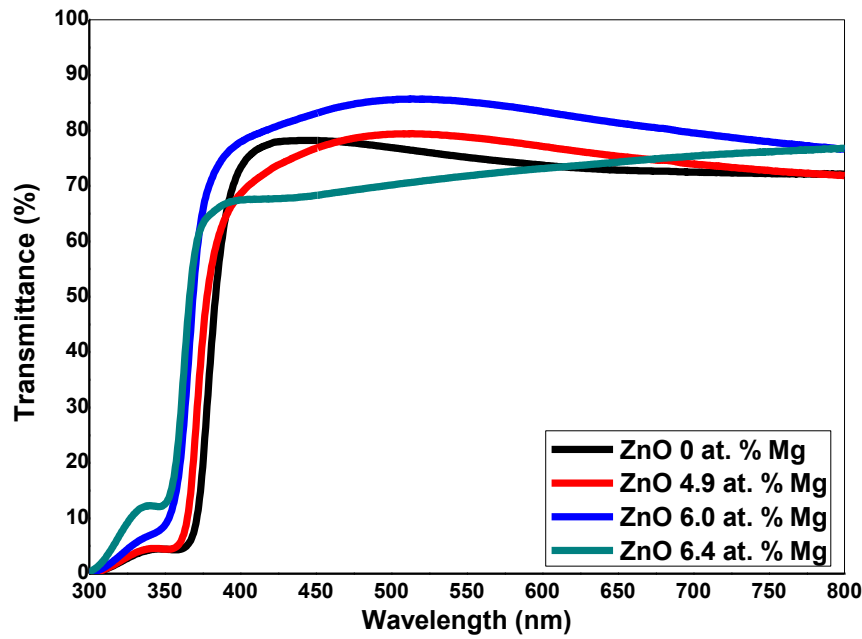




**Fig. 6** ESEM images of: (a) undoped ZnO, (b) 4.9 at. % Mg-ZnO, (c) 6 at. % Mg-ZnO and (d) 6.4 at. % Mg-ZnO, (e) **Cross section** of thin films.

### 3.5 UV- Visible analysis

The optical transmittance of undoped and Mg-doped ZnO films were recorded in 300-800 nm wave-length range and reported in **Figure. 7**. It was noticed that all films exhibited a high level average transparence in the visible range (70-80%) and showed a strong absorption in the UV region; the transmittances edge slightly shifted towards lower wave-length is spotted for the transmittance edge of Mg-doped ZnO films. This is due to an increase in the band-gap energy of Mg-doped ZnO.

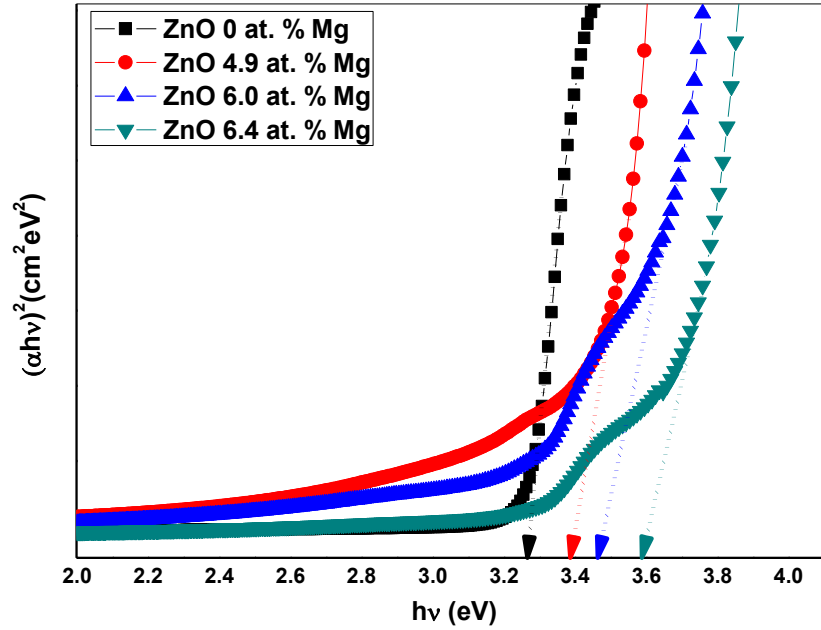


**Fig. 7** Transmittance spectra of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films.

The optical band-gap was determined from the transmittance spectra using the Tauc formula [32]

$$(\alpha h\nu)^2 = B(h\nu - E_g) \quad (10)$$

Where  $\alpha$  is the absorption coefficient, B is the constant, h is the Planck's constant,  $\nu$  is the photon frequency, and  $E_g$  is the optical band-gap. When  $(\alpha h\nu)^2 = 0$ ,  $E_g = h\nu$ , an extrapolation of the linear region of the plot of  $(\alpha h\nu)^2$  versus photon energy ( $h\nu$ ) on the x-axis gives the value of the optical band gap  $E_g$ , **Figure. 8** shows the extrapolation of the curves  $(\alpha h\nu)^2$  as a function of the photon energy  $h\nu$ . The band-gap values obtained from undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films were presented in **Table 4**.



**Fig. 8** The extrapolation of the curves  $(\alpha h\nu)^2$  as a function of photon energy  $h\nu$  of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films.

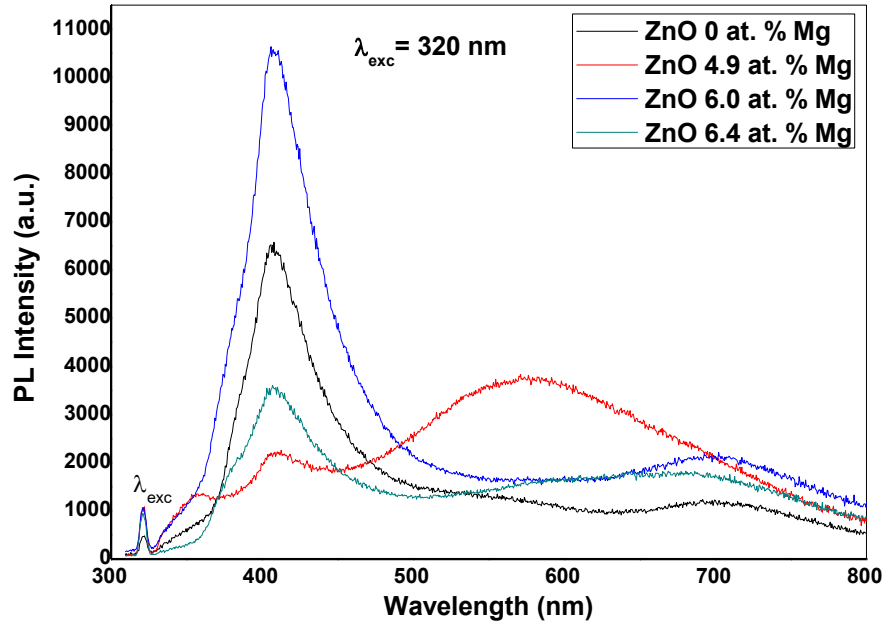
It was observed that the Mg dopant increases the band-gap and this is all the more pronounced as the Mg is increased. This monotonous increase of the band-gap tends to confirm that Mg was well incorporated into the ZnO crystal lattice. The highest band-gap of about 3.59 eV was acquired for 6.4 at. % Mg-doped ZnO as presented in **Table 4**. The increase in band-gap can be linked to the fact that new defects are inserted after  $\text{Mg}^{2+}$  substitute for  $\text{Zn}^{2+}$ , owing to the electronegativity and ionic radius differences between Mg and Zn, there are more electrons shared by Mg dopant related to the lower electron affinity of MgO compared to ZnO. This causes a shift of the Fermi level inside the conduction band in the case of Mg-doped ZnO, or because of the wide band-gap of MgO comparing to ZnO [6, 16]. These results make our deposited films candidates for use in optoelectronics, photovoltaics and multiple optical applications although this is not the focus of the present study.

**Table 4** Band-gap of undoped and Mg-doped ZnO films obtained by UV-Visible spectroscopy.

Samples	Band-gap (eV)
Undoped ZnO	3.28
Mg-ZnO (4.9 at. %)	3.39
Mg-ZnO (6 at. %)	3.48
Mg-ZnO (6.4 at. %)	3.59

### 3.6 Photoluminescence (PL) analysis

The PL measurements of undoped and Mg-doped ZnO films were carried out at room temperature (excitation wave-length 325 nm) in the range of 300-800 nm. The photoluminescence spectra of undoped and Mg-doped ZnO films were shown in **Figure. 9**. Emission peaks of undoped and Mg-doped ZnO exhibit a near band edge emission (NBE) (UV PL) and violet- yellow- red emission (visible PL) (**Figure. 9**). To define the center of every luminescence band, the signal has been deconvoluted with Gaussian curves. The deconvoluted photoluminescence of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films can be observed in **Figure. 10**.

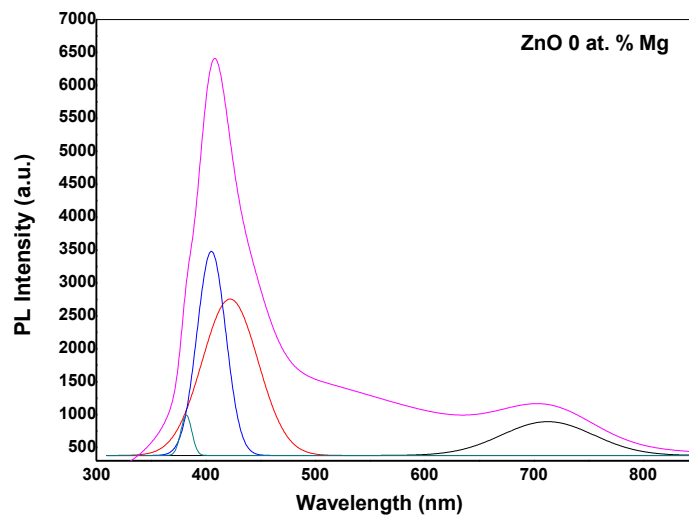


**Fig. 9** Photoluminescence spectra of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films.

The emission spectra show two main emissions. One is the near band edge emission (NBE) at the boundary between ultraviolet and visible located at 364 nm

(3.40 eV), 377 nm (3.28 eV) and 382 nm (3.24 eV), depending on the samples, originated from free-exciton recombination in the UV region. Its position blue-shifts from 382 nm (3.24 eV) to 364 nm (3.40 eV) with Magnesium addition, all the more as the Mg dopant are increased. This is in line with the results gained from spectrophotometry. Hence, the band-gap of ZnO is controlled by the Mg content. This may be due to the impurities, defects and crystal quality [33].

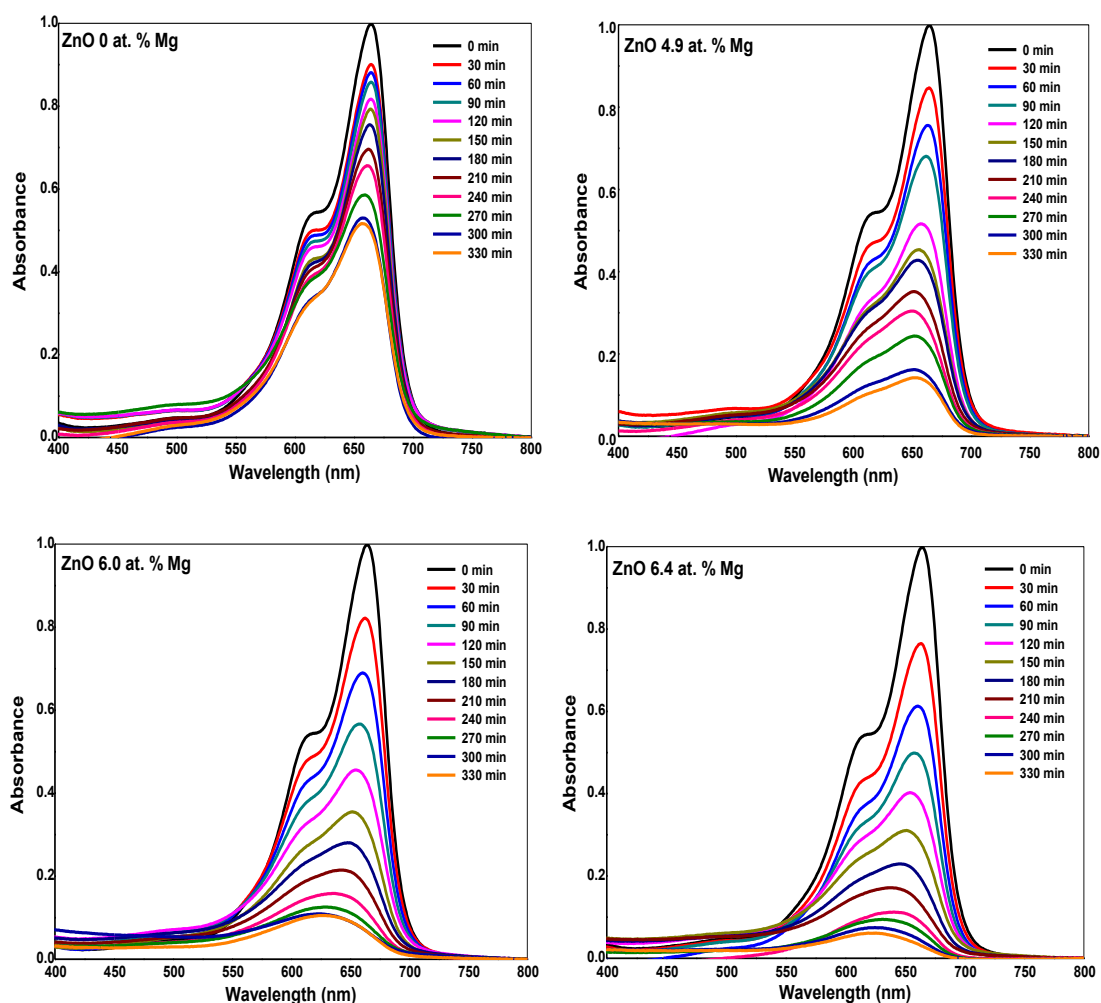
The second one is the broad visible emission located at 405 nm (3.06 eV), 407 nm (3.03 eV), 408 nm (3.03 eV), 416 nm (2.98 eV), 422 nm (2.93 eV), 446 nm (2.78 eV), 478 nm (2.59 eV), 584 nm (2.12 eV), 631 nm (1.96 eV), 690 nm (1.79 eV), 700 nm (1.77 eV) and 712 nm (1.74 eV), depending on sample. It is known to be composed of various contributions usually ascribed to deep level defects in the visible range. Different types of impurities and structural defects such as zinc interstitials (Zni), zinc vacancies (VZn), oxygen interstitials (Oi), antisite oxygen (OZn), zinc antisite (ZnO) and donor-acceptor pair (DAP) have been proposed as the origins of this deep-level emission (DLE) broad band [34, 35, 36].





### 3.7 Photocatalytic activity

Photocatalytic degradation of the synthesized undoped and Mg-doped ZnO films was carried out using the photo-degradation process of organic MB dye under UV-light (365 nm), the photolysis of MB and catalyst the dark adsorption was done before realizing the photocatalytic. **Figure. 11** displays the absorbance spectra of MB as a function of irradiation time of undoped and Mg-doped ZnO films. The curve of MB contains two absorption peaks in the domain of 400 to 800 nm; the absorption peak at 664 nm was selected as the reference to examine the degradation abilities of undoped and Mg-doped ZnO. The intensity of main absorption peaks decreases gradually with light exposure as a function of time using undoped and Mg-doped ZnO films. After 330 min, Mg-doped ZnO (Mg: 6 and 6.4 at. %) enables us to completely degrade MB in water.



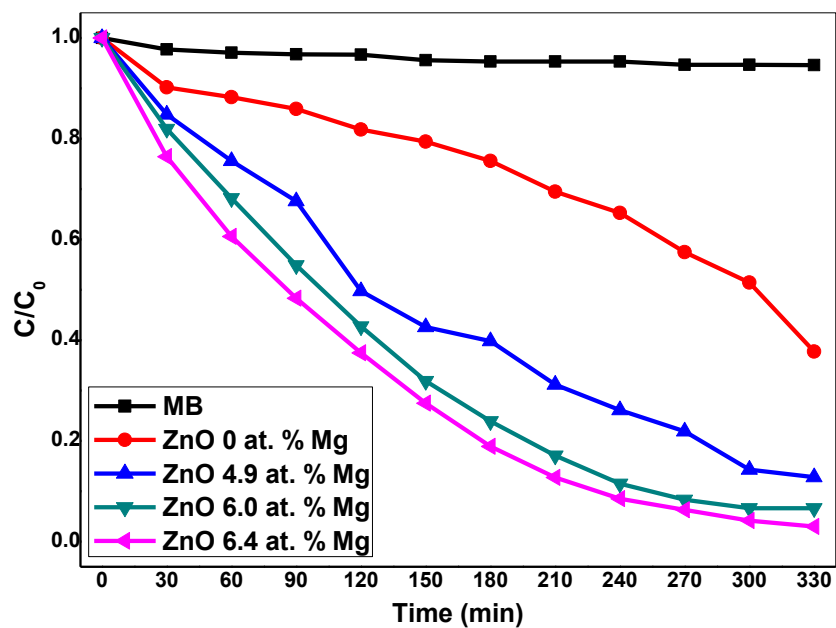
**Fig. 11** Absorbance spectra of MB versus irradiation time of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films as a function of wavelength.

**Figure. 12** shows the concentration ratio ( $C/C_0$ ) of MB dye versus the time of undoped and Mg-doped ZnO films. It is clear that there is no degradation of MB in the absence of undoped and Mg-doped ZnO under irradiation. Obviously, the photocatalytic performance of Mg-doped ZnO is greatly improved and the 6.4 at. % Mg-doped ZnO has the best degradation efficiency and exhibited faster degradation. The kinetics of the reaction was estimated by plotting the natural logarithm of the concentration ratios of MB versus the time of irradiation, as shown in **Figure. 13**. The straight lines showing that the reactions are of pseudo-first order and the apparent rate constants were predestined as the slope of the linear regression to the data points in the plots and the values were listed in **Table 5**. It can be found that the reaction rate has significantly improved after doping modification compared with ZnO ( $k=1.97 \times 10^{-3} \text{ min}^{-1}$ ). Among them, ZnO 6.4 at. % Mg has the highest reaction rate ( $k=8.65 \times 10^{-3} \text{ min}^{-1}$ ), which confirms the faster removal of MB in water compared to ZnO. The relationship between the degradation efficiency of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films and time are spotted in **Figure. 14**. From these curves, it is evident that the degradation efficiency achieved 62%, 87%, 93%, and 97% for Mg 0, 4.9, 6 and 6.4 at. %, respectively.

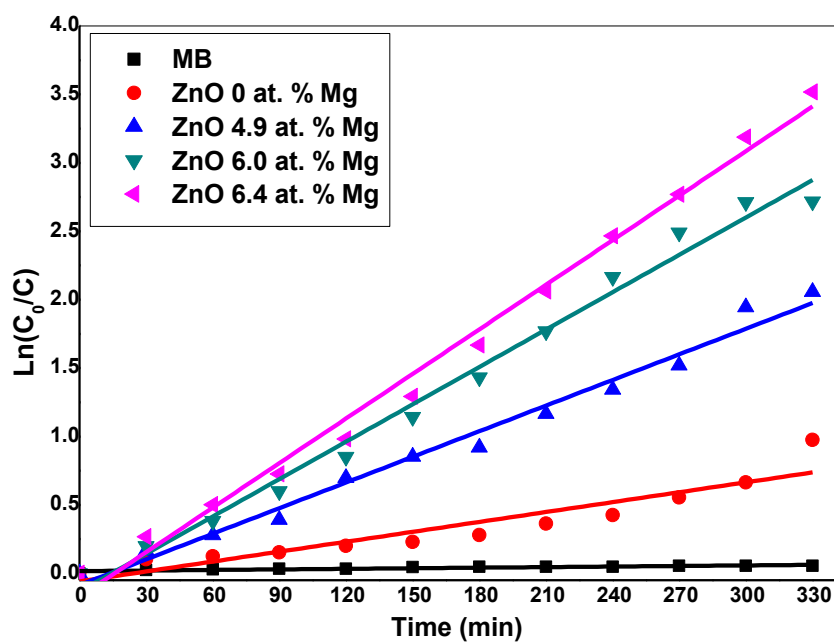
**Table 5** Variation of  $k_{\text{app}}$  values as a function of Mg doping.

Samples	$k_{\text{app}} * 10^{-3} (\text{min}^{-1})$
Undoped ZnO	1.97
Mg-ZnO (4.9 at. %)	5.00
Mg-ZnO (6 at. %)	7.23
Mg-ZnO (6.4 at. %)	8.65

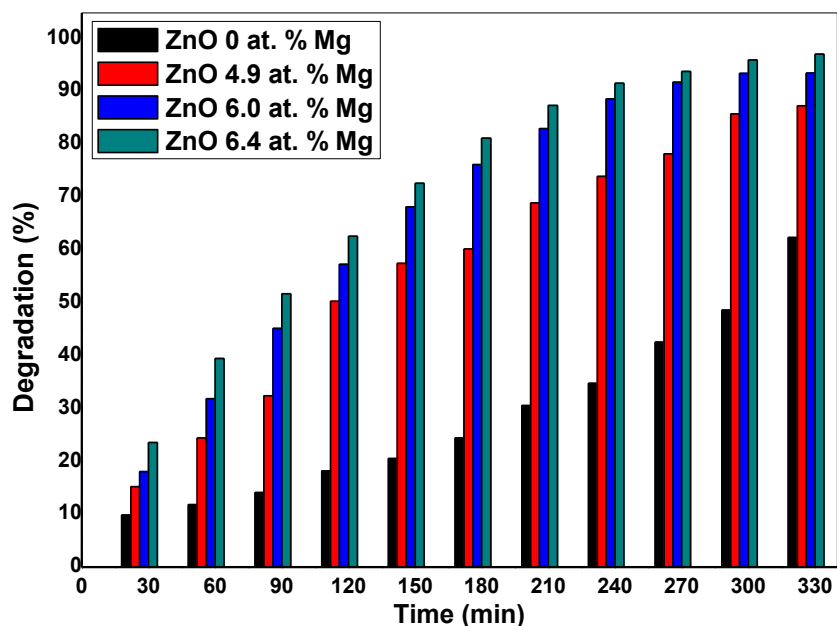




**Fig. 12** the effect of the Mg dopant on the MB dye solution under irradiation of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films as a function of time.



**Fig. 13**  $\ln(C_0/C)$  vs time curves of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films as a function of time.

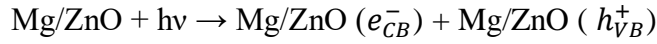


**Fig. 14** Degradation efficiency of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films as a function of time.

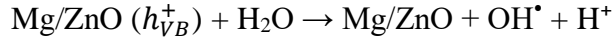
The photocatalytic process is accelerated with an increase in magnesium dopant. Mg-doped ZnO samples display a much higher photocatalytic than undoped ZnO, high photocatalytic degradation efficiency is observed for 6.4 at. % Mg-doped ZnO. The improvement of the photocatalytic performance of ZnO due to Mg doping could be attributed to a number of factors. First, an increase in the band-gap due to Mg dopant causes the higher redox potential of the photo-generated electron-hole pairs, which increases the photocatalytic efficiency of the Mg-doped ZnO, the recombination time of electron-hole, is appropriately delayed due to the doping. Thereby, more hydroxyl radicals can be produced [16]. Benzitouni et al [37] have also reported that the increase in the band-gap of ZnO improve the photocatalytic due to the higher redox potential of the photo-excited pairs electron-hole.

Furthermore, the rougher surface of Mg-doped ZnO, exhibiting a larger specific surface area, could produce more active sites for reactant molecules to promote the efficiency of the electron-hole separation, and for the adsorption of pollutant molecules, therefore producing a higher photocatalytic [38], Jongnavakit et al. [39] have reported that the roughness surface relates to the surface area, which is an important factor that can ameliorate the photocatalytic, the photocatalytic of the ZnO increases with the roughness of the films.

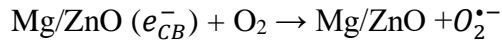
The equations below explained the decolorization mechanism [37]:



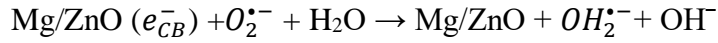
Holes react with surface bound hydroxyl groups  $\text{HO}^-$  or with water  $\text{H}_2\text{O}$  adsorbed on the ZnO surface to constitute hydroxyl radicals  $\text{HO}^\bullet$  and hydrogen ions  $\text{H}^+$ , according to the equations below:



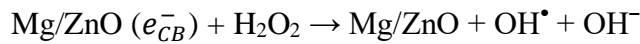
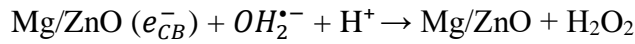
Producing superoxide anion  $\text{O}_2^{\bullet-}$  by electrons are transferred to be adsorbed by oxygen  $\text{O}_2$



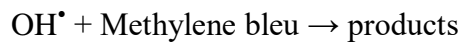
These superoxide anions  $\text{O}_2^{\bullet-}$  additionally react with adsorbed water molecules forming peroxide radicals  $\text{OH}_2^{\bullet-}$  and hydroxyl ions  $\text{OH}^-$ .



The peroxide radicals  $\text{OH}_2^{\bullet-}$  combine with  $\text{H}^+$  producing in the formation of hydroxyl radicals  $\text{OH}^\bullet$  and hydroxyl ions  $\text{OH}^-$ . Hydrogen peroxide is produced as an intermediate product.



Holes oxidize these hydroxyl ions  $\text{OH}^-$  to hydroxyl radicals  $\text{OH}^\bullet$ . All the species facilitate the production of  $\text{OH}^\bullet$ . The strong oxidants  $\text{OH}^\bullet$  reacts with adsorbed dye molecules and perform degradation according to the reaction below:



The photodegradation phenomenon is concentrated on the dye oxidation/reduction through the created electrons from the semiconductor, under UV-light.

#### 4. Conclusion

In the present work, we have investigated the performance of undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films for the MB degradation in aqueous solution. Undoped and Mg-doped ZnO (Mg: 4.9, 6 and 6.4 at. %) films were fabricated by sol-gel process associated with dip-coating. **The EDSX was confirmed the formation of ZnO and the presence of Mg dopant in ZnO.** According to structural studies, all samples exhibited the hexagonal wurtzite structure of ZnO; the addition of Mg improves the film crystallinity and promotes the growth along the c axis of the wurtzite cell. The visible optical transmittance of all the films is between 70 and 80% in the visible range, the band-gap was increased from 3.28 eV to 3.59 eV with Mg dopant. AFM images showed that the undoped and Mg-doped ZnO are composed of rounded shaped particles, the RMS roughness values increases with the increase of Mg dopant from 4.46 to 12.30 nm. ESEM images showed that the undoped and Mg-doped ZnO films had wrinkles morphology. PL measurements confirm the increase of the band-gap energy upon Mg dopant. The catalytic efficiency of ZnO films was improved with the Mg dopant, 97% MB was degraded after 330 min exposure to UV-light.

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