



HAL
open science

Thermal ageing of a silane-crosslinked polyethylene stabilised with an excess of Irganox 1076

Anne Xu, Sébastien Roland, Xavier Colin

► **To cite this version:**

Anne Xu, Sébastien Roland, Xavier Colin. Thermal ageing of a silane-crosslinked polyethylene stabilised with an excess of Irganox 1076 . *Polymer Degradation and Stability*, 2021, 189, pp.109597. 10.1016/j.polymdegradstab.2021.109597 . hal-03262685

HAL Id: hal-03262685

<https://hal.science/hal-03262685>

Submitted on 16 Jun 2021

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Thermal ageing of a silane-crosslinked polyethylene stabilised with an excess of Irganox 1076®

Anne Xu, Sébastien Roland, Xavier Colin*

PIMM, Arts et Métiers Institute of Technology, CNRS, CNAM, HESAM University, 151 boulevard de l'Hôpital, 75013 Paris, France

A B S T R A C T

This work focuses on the thermal ageing in air of a silane-crosslinked polyethylene stabilised with an excess of phenolic antioxidant (Irganox 1076®) in the temperature range between 87°C and 130°C. In such case, it was shown in previous works that the antioxidant is present in two physical states: dissolved in the polymer matrix and exuded at the sample surface under the form of crystals. The purpose was, in particular, to investigate the role of an antioxidant excess in the stabilisation process of the polymer. On this purpose, the pure antioxidant crystals (i.e. the commercial powder) and a polymer sample stabilised with only dissolved antioxidants (i.e. in a concentration lower than the solubility threshold) were investigated at 130°C to isolate and elucidate the thermal ageing behaviour of each antioxidant phase. For all the three studied samples, quinone methide and cinnamate species were formed through the chemical consumption of antioxidants. For both stabilised samples (with an excess, or not, of antioxidants), similar behaviours were observed in term of antioxidant depletion. In particular, it was shown that although some chemical consumption was also detected, the initial antioxidant depletion was essentially due to its physical loss through evaporation. In addition, in both cases, the polymer oxidation was observed after the almost total depletion of antioxidants and without any additional oxidation induction period. The comparison of the oxidation induction periods obtained at 130°C for these two samples showed that the antioxidant excess efficiently participates to polymer stabilisation during thermal ageing, which could be due to a further solubilisation of antioxidants during thermal exposure.

Keywords:

Silane-crosslinked polyethylene
thermal ageing
stabilisation
Irganox 1076®
phenolic antioxidant
heterogeneous distribution
physical loss
exudation
chemical consumption

1. Introduction

Crosslinked low density polyethylene (XLPE) is one of the most widely used polymers for the insulation of electrical cables in the reactor buildings of nuclear power plants (NPP). In this type of application, XLPE is subjected to the combined effects of irradiation and temperature. Under normal service conditions, the dose rate is typically ranged between 10^{-2} and 10^{-1} Gy.h⁻¹ and the temperature is about 30-50°C. However, under severe accidental conditions, e.g. in the event of a loss of primary coolant accident, the dose rate and the temperature can rise up to 10^5 Gy.h⁻¹ and 150°C, respectively [1]. In order to guarantee the control and safety of NPP under both normal and accidental conditions, it is hence crucial to check the long-term resistance to radio-thermal ageing of XLPE insulation.

As any hydrocarbon polymer, XLPE is sensitive to radical chain oxidation which, in a nuclear environment, can be initiated both by ionizing radiations [2-5] and temperature [6-10]. In fact, oxidative

degradation is susceptible to occur at all stages of the polymer life-cycle (from its synthesis up to its in-service use, including its processing) and results into irreversible modifications of the polymer chemical composition, leading, more or less directly, to a loss of the polymer macroscopic properties. In particular, it is often considered that the loss of insulating properties is due both to the formation of more polar species (i.e. oxidation products) and the increase in water absorption [11,12]. In contrast, mechanical embrittlement results essentially from chain scissions and chemically crystallisation phenomenon [13,14].

In order to protect the polymers against oxidative degradation and ensure their long-term stability, a low concentration of antioxidants (AO) is generally incorporated into polymers during the first steps of their processing operation. Different types of AO exist, which mainly differ from their general chemical structure and their stabilisation mechanism. Two main types of AO are commonly used for NPP applications [15-17]: free radicals scavengers (also often called primary AO), and hydroperoxides decomposers. Among the primary AO, hindered phenols, such as octadecyl-3-(3,5-di-tert-butyl-4-hydroxyphenyl)propionate, well-known as Irganox 1076®, have long been known to be effective against thermal oxidation in

* Corresponding author.

E-mail address: xavier.colin@ensam.eu (X. Colin).

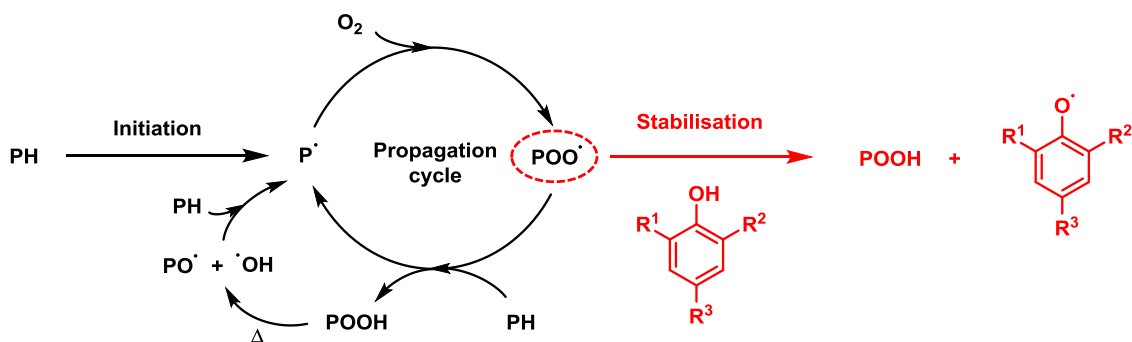


Fig. 1. Simplified mechanistic scheme for the radio-thermal oxidation of a polyolefin stabilised by a hindered phenol antioxidant [18,23–26].

polyolefins in a large temperature range [18–22]. Their stabilisation mechanism stems from the H-donor character of the phenol function of AO, i.e. its great facility to give up its H atom to a free radical, which induces the termination of peroxy radicals (POO^\bullet) generated by the polymer oxidation reaction [18,23–26], as shown in Fig. 1.

However, due to their significant difference in polarity, the solubility of hindered phenols in polyolefins is generally low, typically less than 1 wt% in polyethylene (PE) [17,27,28]. Due to this limited solubility, an oversaturation can easily occur in industrially formulated polyolefins. In fact, AO are generally incorporated into polymers during the processing operation at high temperatures in melt state, i.e. in conditions where AO are usually highly soluble in polymers. However, during cooling, as the polymer crystallises and the AO solubility in the amorphous phase decreases with temperature, the polymer can rapidly become supersaturated in AO, leading to several phenomena such as a phase separation of AO, leading to the exudation and blooming of AO at the sample surface [29–32].

In our previous studies on a PE-based material stabilised with different concentrations of Irganox 1076® [29,33], a phase separation was detected above a threshold concentration of AO of about 0.9 wt% (i.e. $\approx 1.6 \times 10^{-2} \text{ mol.L}^{-1}$) at ambient temperature. In particular, it was shown that AO in excess were under the form of crystals and only present onto the sample surface. Thus, two different situations were distinguished depending on the AO concentration. When the AO concentration $[\text{AO}]$ is below or equal to the threshold concentration (noted $[\text{AO}]_c$), the stabilisation is considered homogeneous because all AO are dissolved in the polymer matrix. In contrast, when $[\text{AO}]$ is above $[\text{AO}]_c$, the stabilisation is considered heterogeneous because the polymer matrix is fully saturated with dissolved AO but, in addition, AO in excess are exuded and crystallised at the sample surface. The question that naturally arises is: do AO crystals also contribute to the polymer stabilisation under the different exposure conditions of a nuclear environment?

The present study aims at investigating the stabilisation efficiency by Irganox 1076® in a heterogeneously stabilised silane-crosslinked polyethylene (Si-XLPE) during its thermal ageing in air at three different temperatures: 87°C, 110°C, and 130°C. The analyses will be performed using the same experimental methodology established in our previous publications [29,33]. Given the microstructural complexity of this material, the first stage will consist in isolating and elucidating the thermal ageing behaviour of its two constitutive phases (i.e. AO crystals and homogeneously stabilised Si-XLPE matrix) by additionally examining three model materials in turn (as schematically represented in Fig. 2): i) the as-received Irganox 1076® powder; ii) a Si-XLPE sample stabilised with 0.6 wt% of Irganox 1076® (i.e. below the threshold concentration of 0.9 wt%); iii) the pure Si-XLPE matrix. This first stage is necessary not only to determine the ageing mechanisms in each phase, but

also to deduce the possible additional ageing mechanisms operating in the more complex material under study. Then, the study of the thermal ageing of the heterogeneously stabilised sample will enable to check the reliability of the previous ageing mechanisms in the more complex material, and to investigate the stabilisation efficiency by antioxidants in excess (crystals).

2. Materials and methods

2.1. KBr disc with Irganox 1076®

The commercial Irganox 1076® (whose chemical structure is shown in Fig. 3) was supplied by Sigma Aldrich under the form of a white powder. It should be mentioned that some physico-chemical characteristics of the as-received commercial powder were already determined and discussed in our previous work [29]. In particular, it was shown that its crystalline structure (i.e. form I) is the same than the AO crystals exuded at the surface of the heterogeneously stabilised samples.

The thin KBr disc was prepared by mixing 100 mg of KBr powder with 1 mg of Irganox 1076® powder. The resulting mixture was grinded into a fine powder using a mortar then compressed for 5 minutes under 4 MPa to obtain the KBr disc.

2.2. Films of silane-crosslinked polyethylene stabilised with Irganox 1076®

Films of silane-crosslinked polyethylene (Si-XLPE) of about 600 μm thick containing about 0.6 wt% and 2.1 wt% of Irganox 1076® were gratefully provided by Nexans. These films were respectively produced by hot-compression and extrusion moulding from pellets of a linear low-density polyethylene grafted with a vinyl trimethoxy silane (Si-g-LLDPE). The crosslinking of the polymer was performed by immersion in water at 65°C for 48 hours. The density, crystallinity ratio and gel content of the resulting Si-XLPE are about 0.918 g.cm^{-3} , 41%, and 71%, respectively.

It should be mentioned that the pellets of Si-g-LLDPE already contained a small amount of storage AO (about 0.1 wt%), mainly BHT and Irganox 1076®, prior to the addition of Irganox 1076® during the processing operation [29]. These storage AO were considered for the calculation of the total concentration of AO in the Si-XLPE films, thus inducing a slight difference between the total concentrations of ester and phenol functions.

All films were stored at ambient temperature in a desiccator containing silica gel before any physico-chemical characterization.

2.3. Surface cleaning treatment

The surface of the heterogeneously stabilised Si-XLPE films were gently cleaned with absorbing paper and acetone as polar

Additional materials

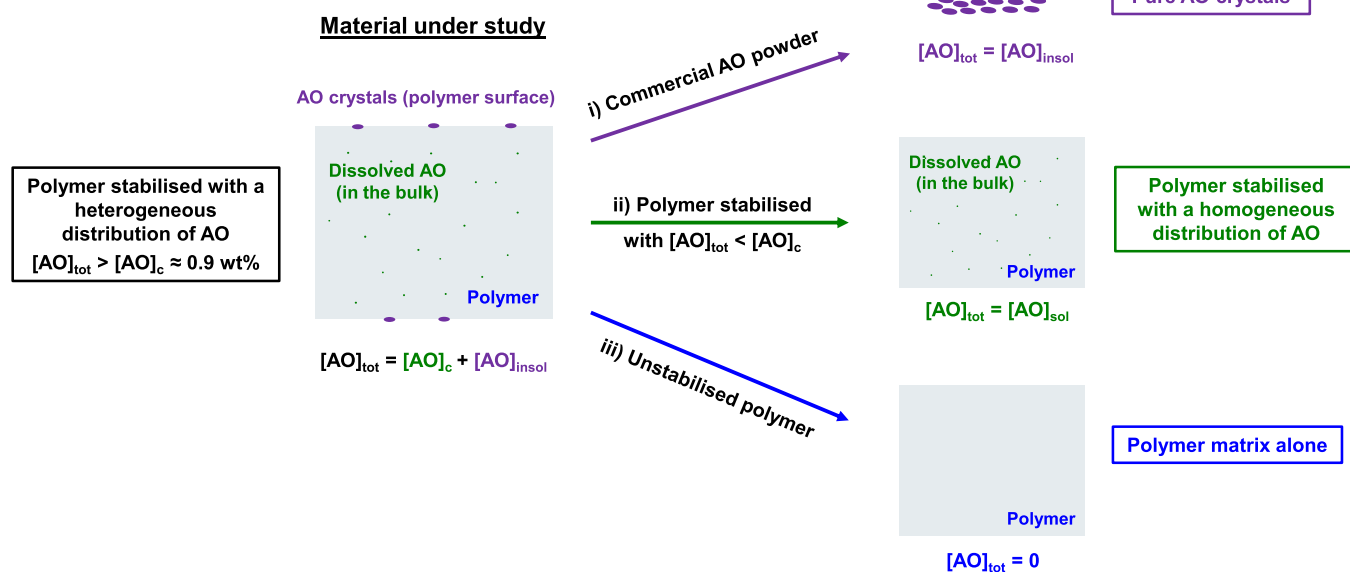


Fig. 2. Schematic representation of the microstructure of the heterogeneously stabilised polymer under study and the three model materials to be additionally investigated in order to elucidate the ageing mechanisms. $[AO]_c$ corresponds to the solubility threshold of AO in the polymer matrix.

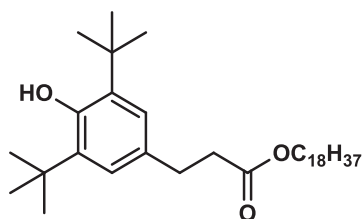


Fig. 3. Chemical structure of Irganox 1076®.

solvent (supplied by Sigma-Aldrich) in order to remove all AO crystals present on their surface without changing the concentration of the AO dissolved in their bulk, as previously shown in reference [33]. Thanks to this surface cleaning, it was possible to deconvolve the physico-chemical responses of dissolved and crystallised AO.

2.4. Ageing conditions

All the thermal ageing experiments were performed in air-ventilated thermal ovens. The pure Irganox 1076® powder (in KBr disc) and the Si-XLPE film stabilised with 0.6 wt% of Irganox 1076® were only thermally aged at 130°C. In contrast, the Si-XLPE film stabilised with 2.1 wt% of Irganox 1076® was exposed to three different temperatures of ageing: 87°C, 110°C, and 130°C. It should also be mentioned that an almost pure Si-XLPE film, i.e. produced from pellets of Si-g-LLDPE without the addition of Irganox 1076®, was also aged under the same exposure conditions as a control sample.

For each ageing temperature, samples were removed regularly from ovens and cooled to room temperature in the dry atmosphere of a desiccator before being analysed by complementary physico-chemical characterisation techniques.

2.5. Fourier-Transform Infrared (FTIR) spectroscopy

This technique was used to monitor the modifications of the samples chemical composition. In particular, FTIR spectroscopy was

used to identify and quantify the different AO phases (dissolved and crystallised) and the oxidation products (from AO and the polymer matrix) detected during the thermal exposure in air. For these purposes, this technique was used in two different modes: Attenuated Total Reflectance (ATR) mode and transmission mode. In fact, in ATR mode, the penetration depth of the IR radiation in PE samples is generally ranged between 0.2 and 1.3 μm , for a wavenumber typically varying from 4000 to 650 cm^{-1} [29]. Therefore, in ATR mode, mainly the surface of the polymer films is analysed. In contrast, in transmission mode, the sample is analysed throughout its whole thickness (i.e. both the surface and the bulk). With regard to AO, in ATR mode, mainly the IR signature of the AO crystals present on the sample surface is detected, while in transmission mode, both the IR signatures of dissolved and crystallised AO can be observed.

Previous works have indeed shown that FTIR spectroscopy is a very efficient technique to identify the physical state (i.e. dissolved or crystallised) of AO in PE-based materials [29,33], but also to distinguish the different polymorphic forms of AO crystals [29,31] as, in each case, a specific IR signature can be observed by FTIR spectroscopy. The position of the IR absorption bands of interest of AO and their attribution, depending on the physical state or the crystalline structure of AO, are compiled in Table 1.

Table 1
IR absorption bands of Irganox 1076® and their attribution [29,31,33].

Attribution	Wavenumber (cm^{-1})
In "dispersed state" (i.e. dissolved in PE)	3648-3649 (O-H of phenol)
	1740-1741 (C=O of ester)
Crystalline Form I	3639-3642 (O-H of phenol)
	1733-1736 (C=O of ester)
	1279 (C-O of ester)
Crystalline Form II	3590 (O-H of phenol)
	1723-1726 (C=O of ester)
	1288 (C-O of ester)
Crystalline Form III	3609-3612 (O-H of phenol)
	1720-1723 (C=O of ester)
	1288 (C-O of ester)

When performed in transmission mode, FTIR spectroscopy can be used to estimate the concentration of chemical species using the Beer-Lambert's law, which can be written as follows:

$$\mathbf{A} = \boldsymbol{\varepsilon} \times \mathbf{e} \times \mathbf{C} \quad (1)$$

Where \mathbf{A} is the absorbance, \mathbf{e} the sample thickness (cm), \mathbf{C} the concentration of the absorbing species (mol.L^{-1}), and $\boldsymbol{\varepsilon}$ the molar extinction coefficient ($\text{L.mol}^{-1}.\text{cm}^{-1}$). AO in the Si-XLPE films were titrated using the FTIR calibration curves established both for the phenol and ester functions of Irganox 1076® in our previous studies [29,33]. Moreover, AO oxidation products were quantified using molar extinction coefficients from literature. Their values will be given later in the results and discussion section.

FTIR spectra were recorded at room temperature using a Perkin Elmer FTIR Frontier spectrometer from 4000 to 650 cm^{-1} in ATR mode, and from 4000 to 400 cm^{-1} in transmission mode, after averaging the 16 scans obtained with a minimum resolution of 4 cm^{-1} . In the case of ATR mode, the analyses were performed with a diamond/ZnSe crystal. Moreover, in order to better monitor all the chemical and physical modifications throughout thermal ageing, the FTIR spectra obtained in ATR mode were normalised with the absorbance of the CH_2 scissoring vibrations at 1472 cm^{-1} assigned to the PE crystal phase [34]. In contrast, the FTIR spectra obtained in transmission mode were only normalised with the films thickness.

2.6. Ultraviolet spectroscopy (UV)

This technique was used as a complementary technique of FTIR spectroscopy to monitor the AO depletion and the formation of its main oxidation products during the thermal exposure in air.

Indeed, the phenol function of Irganox 1076® can be detected around 270-280 nm, as reported in the literature [35], while its main oxidation products can be detected around 300-320 nm [36]. As for FTIR spectroscopy, the Beer-Lambert's law (Eq. 1) can also be applied in UV spectroscopy to quantify the chemical species. The main interest of this technique lies in the fact that the molar extinction coefficients for AO and its oxidation products are generally higher in UV than in IR, thus allowing a better detection in UV, even at low concentrations. However, for the same reason, the UV absorption bands of interest can be more easily saturated, in particular when high concentrations are concerned. For the titration by UV spectroscopy, the molar extinction coefficients were taken from literature. Their values will be given later in the results and discussion section.

UV spectra were recorded at room temperature using a Perkin Elmer Lambda 35 UV-Visible spectrometer from 200 to 400 nm with a scanning rate of 60 nm.min^{-1} in transmission mode.

2.7. Differential Scanning Calorimetry under O_2 (OIT measurements)

This technique was used to measure the oxidation induction time (OIT) at a given temperature of the stabilised Si-XLPE films throughout their thermal exposure in air, which gives an estimation of the global concentration of active AO in PE-based materials [33,37]. Such a correlation can be written as follows:

$$\text{OIT}(T) = K_T \times [\text{AO}] + \text{OIT}_0(T) \quad (2)$$

Where K_T is a constant only depending on the temperature T and the couple {polymer matrix, AO} under consideration, $[\text{AO}]$ the concentration of the active functions of AO (here phenols), and $\text{OIT}(T)$ and $\text{OIT}_0(T)$ the values of OIT at the measurement temperature for the stabilised and non-stabilised polymer matrix, respectively. For the pure Si-XLPE matrix, OIT_0 is near zero in pure O_2 atmosphere at 205°C.

Table 2
Melting properties of Irganox 1076® depending on its crystalline structure [29,31,38].

Crystalline form of Irganox 1076®	Form I	Form II	Form III
Melting point (°C)	51 ± 1	51 ± 1	47 ± 1
$\Delta H_{\infty, \text{AO}}$ (J.g^{-1})	120 ± 3	123 ± 4	142 ± 9
$\Delta H_{\infty, \text{AO}}$ (kJ.mol^{-1})	64 ± 2	65 ± 2	75 ± 5

OIT measurements were performed under a pure O_2 flow at 205°C using TA Instrument DSC Q10 and Q20 calorimeters. This temperature was chosen in order to have an acceptable duration of experiments for both the unaged and aged samples. Samples of about 5 mg were introduced in open standard aluminium pans. The samples were first heated under pure N_2 flow (50 mL.min^{-1}) from room temperature to 205°C, at a heating rate of 10°C.min⁻¹. After an isotherm segment of 5 min for temperature equilibration, the gas flow was switched from N_2 to O_2 . An isotherm at 205°C was then performed under pure O_2 flow (50 mL.min^{-1}) in order to access the oxidation induction time (OIT) at this temperature. The OIT value was determined using the tangent method, which corresponds to the duration period between the introduction of the O_2 flow in the DSC cavity and the onset of the oxidation exotherm.

2.8. Differential Scanning Calorimetry under N_2 (DSC measurements)

This technique was used for detecting and quantifying the AO crystals in all samples (i.e. Irganox 1076® powder and Si-XLPE films), but also for monitoring the consequences of oxidation on the crystallinity ratio of the polymer matrix.

In our previous studies [29,33,37], it was shown that AO crystals can also be easily detected by DSC analysis, due to the presence of their melting peak around 50°C on the DSC thermograms. Then, knowing the melting enthalpy of AO crystals, it is possible to estimate the concentration of insoluble AO in the polymer films using the following equation:

$$[\text{AO}]_{\text{insol}} = \frac{\Delta H_{\text{m, AO, polymer}}}{\Delta H_{\infty, \text{AO}}} \times \rho_{\text{PE}} \quad (3)$$

Where ρ_{PE} is the polymer density (0.918 g.cm^{-3}), $\Delta H_{\text{m, AO, polymer}}$ the melting enthalpy of AO crystals in the polymer sample under study (J.g^{-1}), and $\Delta H_{\infty, \text{AO}}$ the melting enthalpy of pure AO crystal (kJ.mol^{-1}). As previously recalled, Irganox 1076® can undergo polymorphism, and each polymorph has specific physico-chemical properties. In particular, the melting properties of the three possible crystalline forms of Irganox 1076® are compiled in Table 2 [29,31,38].

On the other hand, the crystallinity ratio (χ_c) of Si-XLPE can be determined from the area under the corresponding melting peak in the DSC thermogram using the common following equation:

$$\chi_c = \frac{\Delta H_{\text{m}}}{\Delta H_{\infty}} \quad (4)$$

Where ΔH_{m} and ΔH_{∞} are the melting enthalpies (expressed in J.g^{-1}) of the sample under investigation and the PE crystal, respectively. In literature, the commonly used value for ΔH_{∞} is 290 J.g^{-1} [39,40].

DSC thermograms were recorded using a TA instrument DSC Q1000 calorimeter beforehand calibrated with an indium reference. The DSC experiments were performed on samples with a mass ranged between 5 and 8 mg introduced in closed standard aluminium pans. The analyses were performed under a pure N_2 flow of 50 mL.min^{-1} , in the temperature range between -50°C and 250°C, with heating and cooling rates of 10°C.min⁻¹.

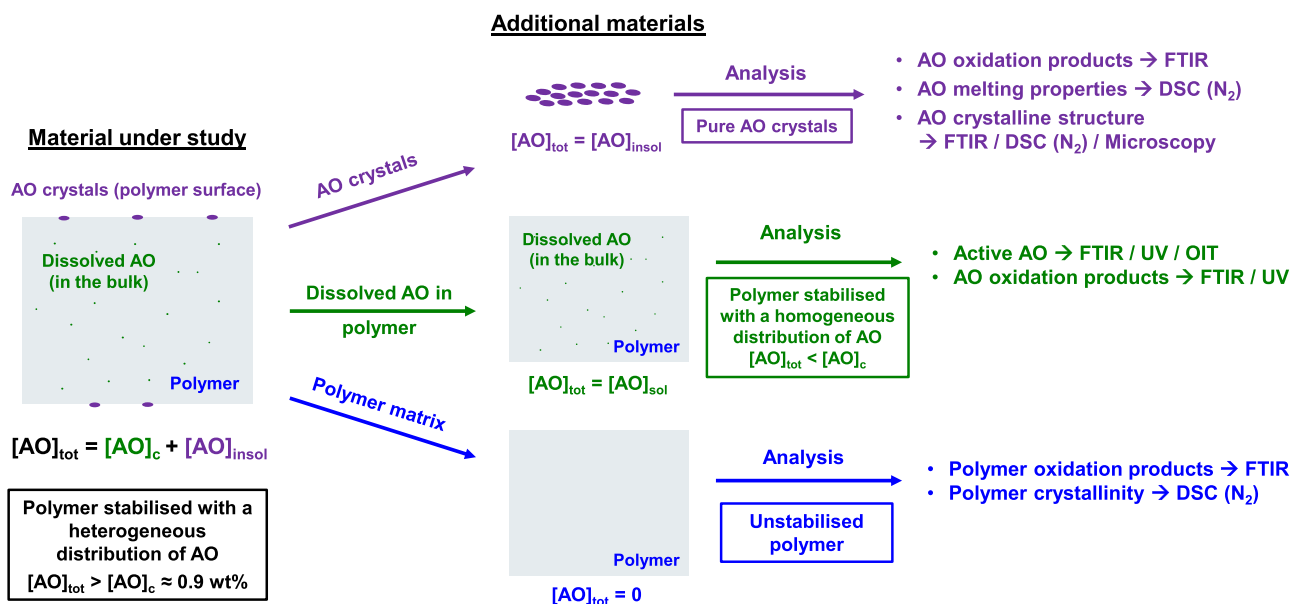


Fig. 4. General experimental methodology used in this work.

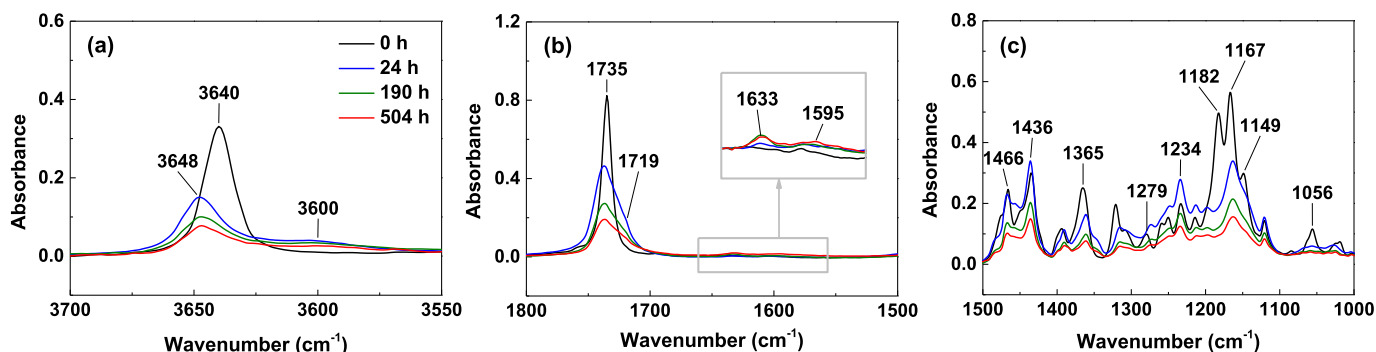


Fig. 5. Changes in the FTIR spectrum in transmission mode of the pure Irganox 1076® powder during its thermal ageing at 130°C in air: (a) between 3700 and 3550 cm⁻¹, (b) between 1800-1500 cm⁻¹, and (c) between 1500-1000 cm⁻¹.

2.9. Optical microscopy

This technique was used to observe the surface of the Si-XLPE films. It is thus complementary to FTIR spectroscopy in ATR mode and DSC analysis under N₂ in order to confirm the presence, or absence, of AO crystals.

The surface of the Si-XLPE films was examined before and after acetone cleaning using a Zeiss Axio Imager 2 optical microscope. The microscopic observations were performed in reflection mode with a polarised light.

The general experimental methodology used for this study is summarised in Fig. 4. It is noteworthy to mention that, in the case of the heterogeneously stabilised sample under study, all the experiments which have just been detailed were used.

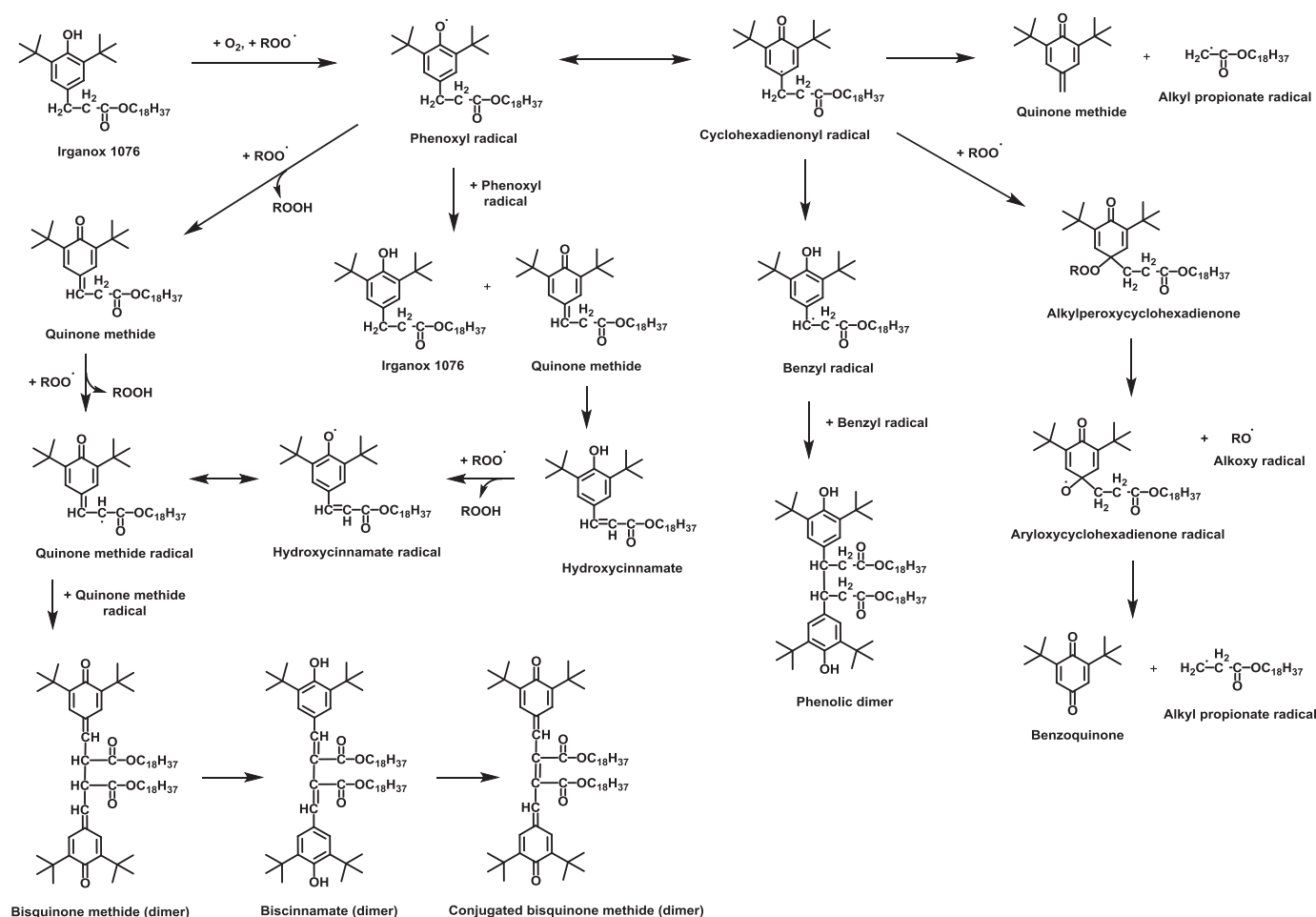
3. Results and discussion

3.1. Thermal ageing at 130°C of pure Irganox 1076®

The KBr disc containing 1 wt% of pure Irganox 1076® powder was aged at 130°C in air. The sampling was regularly performed throughout thermal exposure and the degradation of Irganox 1076® was monitored by FTIR spectroscopy in transmission mode. The FTIR spectra obtained before and after different durations at 130°C in air are shown in Fig. 5.

As previously reminded, Irganox 1076® can exist under different crystalline forms, each of them having a specific IR signature [29,31,38,41]. In our previous work [29], it was shown that the commercially available Irganox 1076® powder was crystallised under the so-called Form I. For this crystalline Form I, the IR bands of the phenol (O-H) and ester functions (C=O) are observed around 3640 cm⁻¹ and 1735 cm⁻¹, respectively, as shown in Fig. 5a and Fig. 5b. After 24 h of exposure at 130°C in air, however, several changes were observed. In the hydroxyls spectral region (i.e. between 3700 and 3550 cm⁻¹, Fig. 5a), the IR band of phenol at 3640 cm⁻¹ seems no longer present, but two other IR bands are now observed at 3648 cm⁻¹ and around 3600 cm⁻¹. Meanwhile, in the carbonyls spectral region (between 1500 and 1800 cm⁻¹, Fig. 5b), the IR band of the ester function of AO at 1735 cm⁻¹ is still present but has significantly decreased, and a broadening band around 1719 cm⁻¹ is observed. In addition, IR bands around 1633 cm⁻¹ and 1600 cm⁻¹ are also appearing. The changes observed in the hydroxyls spectral region could be due to a change in the chemical environment of the phenol functions, either due to a modification of the chemical structure of AO and/or a modification of intermolecular interactions involving the phenol functions of AO.

A first possibility could be a modification of the crystalline structure of AO after thermal exposure at 130°C in air. Let us remember that the as-received Irganox 1076® powder is crystallised



Scheme 1. Reaction pathways proposed for the thermal degradation of pure phenolic AO or polyolefins stabilised with phenolic AO [18–20,42–46].

under the so-called Form I with a melting point around 51°C. In our previous study [29], indeed, it was observed a recrystallization of this AO powder under another crystalline form after a thermal treatment above its melting point, in particular under the Form III which is characterized with a melting point around 47°C. The new positions of the phenol and ester IR bands observed here after 24 h at 130°C could correspond to this crystalline Form III, typically characterized by IR absorptions around 3610 cm^{-1} and 1723 cm^{-1} , respectively [29,31,38]. However, according to Saunier et al. [31], this crystalline Form III has another specific IR band at 1288 cm^{-1} (C–O stretching of the ester). But here, as shown in Fig. 5c, no IR band seems to be detected at 1288 cm^{-1} , which does not allow us to conclude.

Another explanation could involve the change of the chemical structure of AO due to its thermal degradation. Indeed, during the thermal oxidative ageing of the pure phenolic AO, but also of polymers stabilised with phenolic AO, this latter can undergo several chemical transformations, as shown in Scheme 1 [18–20,42–46]. The first radical species formed from phenolic AO during oxidative ageing are phenoxyl radicals, which can be generated by reaction of the phenol function of AO with various dehydrogenation agents, such as peroxy radical (ROO^\cdot) or molecular oxygen (which is a bi-radical) [20]. These phenoxyl radicals can then lead to a long list of transformation products, as shown in Scheme 1. For instance, one of the main degradation products of hindered phenols are quinone methides [18–20,42–46]. According to Pospisil [18], these species can be generated by bimolecular disproportionation, radical attack by peroxy radicals or an intramolecular re-

arrangement. In FTIR spectra, these species can typically be observed around 1630–1640 cm^{-1} (C=O stretching) and 1590–1610 cm^{-1} (C=C stretching) [47]. The small growing IR bands observed around 1633 cm^{-1} and 1600 cm^{-1} in Fig. 5b could thus be attributed to quinone methides. Cinnamate species can also be detected during the oxidative ageing of pure phenolic AO or polymers stabilised with phenolic AO [42,43,45,46]. They can be generated by an intramolecular rearrangement of quinone methides, as shown in Scheme 1. Due to the conjugation effect in unsaturated esters, such as cinnamates, these species are typically observed in FTIR spectra at slightly lower wavenumbers than saturated esters: typically around 1710–1730 cm^{-1} for the C=O stretching of unsaturated esters, against around 1730–1750 cm^{-1} for the C=O stretching of saturated esters [16,48–51]. The IR band observed around 1719 cm^{-1} in Fig. 5b could thus correspond to cinnamate species.

Finally, as ageing in air at 130°C proceeds, most of the IR bands are decreasing with ageing time, at the exception of the IR bands around 1633 and 1600 cm^{-1} , which seem rather to slightly increase, as shown in Fig. 5b. These observations are presumably due to a combination between AO physical loss by evaporation and chemical degradation.

To summarize, several changes were observed by FTIR spectroscopy after the thermal exposure of the pure Irganox 1076® at 130°C in air. Most of the changes were observed just after 24 h of exposure. In particular, the formation of quinone methides (IR bands around 1633 and 1600 cm^{-1}) and cinnamate species (IR band at 1719 cm^{-1}) are suspected to occur. Moreover, although the recrystallization of Irganox 1076® into crystalline Form III (previously

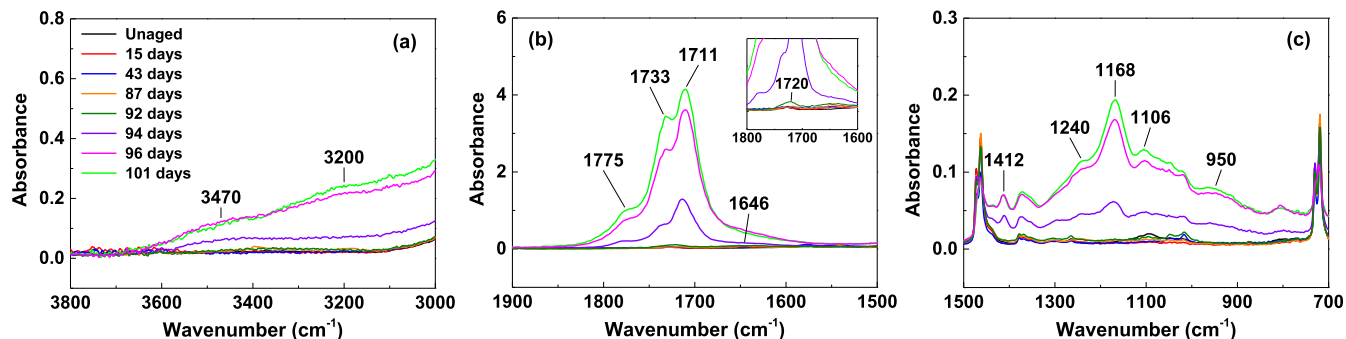


Fig. 6. Changes in the three normalised regions of interest of the FTIR spectrum in ATR mode of the Si-XLPE stabilised with 0.6 wt% of Irganox 1076® during its thermal ageing at 130°C in air: (a) between 3800-3000 cm^{-1} , (b) between 1900-1500 cm^{-1} , and (c) between 1500-700 cm^{-1} .

detected by DSC analysis) was not confirmed by FTIR spectroscopy, a modification of intermolecular interactions (such as H-bonding) involving phenol functions cannot be totally excluded, in particular to explain the formation of two new IR bands in the hydroxyls spectral region (Fig. 5a). Indeed, as new chemical species are formed during the thermal exposure of AO, its chemical environment is modified and new intermolecular interactions can hence occur.

3.2. Physico-chemical characterization of the thermal ageing at 130°C of the Si-XLPE stabilised with a homogeneous distribution of Irganox 1076®

The thermal ageing under air at 130°C of a Si-XLPE stabilised with 0.6 wt% of Irganox 1076® was investigated using the general methodology detailed in Fig. 4. It should be emphasized that, in this sample, the concentration of AO is below the concentration threshold of 0.9 wt%, meaning that all the AO is dissolved in the polymer matrix.

3.2.1. Surface analysis (FTIR spectroscopy in ATR mode)

The FTIR spectra obtained in ATR mode before and after different durations at 130°C in air are shown in Fig. 6. The three spectral regions of interest were normalised with the absorbance of the IR band at 1472 cm^{-1} assigned to the PE crystal phase [31] to better visualize the changes.

As expected, no IR band of Irganox 1076® is initially detected, and only IR bands of Si-XLPE are actually observed at the sample surface. As the ageing proceeds, no significant changes are first detected in the FTIR spectra. Then, after 92 days of exposure at 130°C, OIT at 205°C becomes relatively low (about 3 min), meaning that active AO are still present but in a very low concentration. Therefore, a small IR band, which could be due to ketones from the polymer oxidation, starts to be detected around 1720 cm^{-1} . After 94 days of exposure at 130°C, OIT is null and IR bands corresponding to a wide variety of oxidation products of Si-XLPE are more clearly detected. The large IR band between 3000 and 3500 cm^{-1} is attributed to hydroxyl functions of alcohols, hydroperoxides, and/or carboxylic acids. In the carbonyls region, the main IR bands at 1711 cm^{-1} , 1733 cm^{-1} , and 1775 cm^{-1} correspond to carboxylic acids, esters, and lactones, respectively. The IR signatures of carbonyl and hydroxyl species are also observed between 1000 and 1500 cm^{-1} , in particular around 1410 cm^{-1} and 1240 cm^{-1} where the IR bands are attributed to the C-O stretching of alcohols and/or carboxylic acids, but also around 1169 cm^{-1} and 1110 cm^{-1} where the IR bands are assigned to the C-O stretching of carbonyls. In addition, around 950 cm^{-1} , trans vinylene species (C=C), also resulting from Si-XLPE oxidation, are detected.

3.2.2. Morphological changes (DSC analysis)

The DSC thermograms and the corresponding values of crystallinity ratio of the polymer (determined using Eq. 4) determined before and after different durations at 130°C in air are shown in Fig. 7.

As it can be observed in Fig. 7a, no melting peak of Irganox 1076® around 50°C is detected, confirming that no Irganox 1076® crystal is present in this sample. This observation is in perfect agreement with the previous results of FTIR spectroscopy in ATR mode. Moreover, a modification of the general shape of the melting peak of Si-XLPE can first be pointed out after 15 days of exposure at 130°C. In fact, as the ageing temperature is above the melting point of Si-XLPE (about 114°C), all the polymer crystals are melted during thermal exposure and thus, a total recrystallization occurs during the sample cooling at room temperature before DSC analysis. In addition, as it can be observed in Fig. 7b, a slight decrease in the crystallinity ratio of Si-XLPE is detected during the first ageing duration, which is likely due to a post-crosslinking phenomenon of the polymer matrix. Indeed, such a phenomenon was already evidenced during the thermal ageing at 130°C in air of a Si-XLPE stabilised with a thiopropionate AO [37]. Then, as the thermal ageing proceeds, no significant modification of the crystallinity ratio of Si-XLPE is observed until about 96 days. At the end of this induction period, a sudden and sharp decrease in the crystallinity ratio of Si-XLPE is detected, clearly indicating the beginning of the polymer oxidation, as previously shown by FTIR spectroscopy in ATR mode.

Finally, the formation of a broad exothermic peak between 150 and 220°C is also observed in the DSC thermograms, in particular for the sample aged 94 days at 130°C in air. This exothermic peak is, in fact, due to the thermal decomposition of hydroperoxide species (POOH), which, as shown in Scheme 1, are one of the by-products of the possible stabilisation reactions of the phenolic AO with peroxy radicals. The concentration of POOH can be estimated by DSC from the area under the exothermic peak of their thermal decomposition using the following equation [7,52,53]:

$$[\text{POOH}] = \frac{\Delta H_{\text{POOH}}}{\Delta H_{\text{theory}}} \times \rho_{\text{PE}} \quad (5)$$

Where ρ_{PE} is the polymer density (here 0.918 $\text{g}\cdot\text{cm}^{-3}$) and ΔH_{theory} the theoretical enthalpy of the thermal decomposition of one mole of POOH ($\Delta H_{\text{theory}} = -291 \text{ kJ}\cdot\text{mol}^{-1}$ [52]).

The changes with time of exposure in the POOH concentration (determined by DSC) and the normalised absorbance of carbonyl species due to polymer oxidation (determined by FTIR spectroscopy in ATR mode at 1711 cm^{-1}) are compared in Fig. 8. As previously said, POOH are detected after 94 days of exposure at 130°C in air with a maximum concentration of about 0.3 $\text{mol}\cdot\text{L}^{-1}$. It can be observed that this duration exactly corresponds to the induction time of carbonyl species. At this time of exposure, all

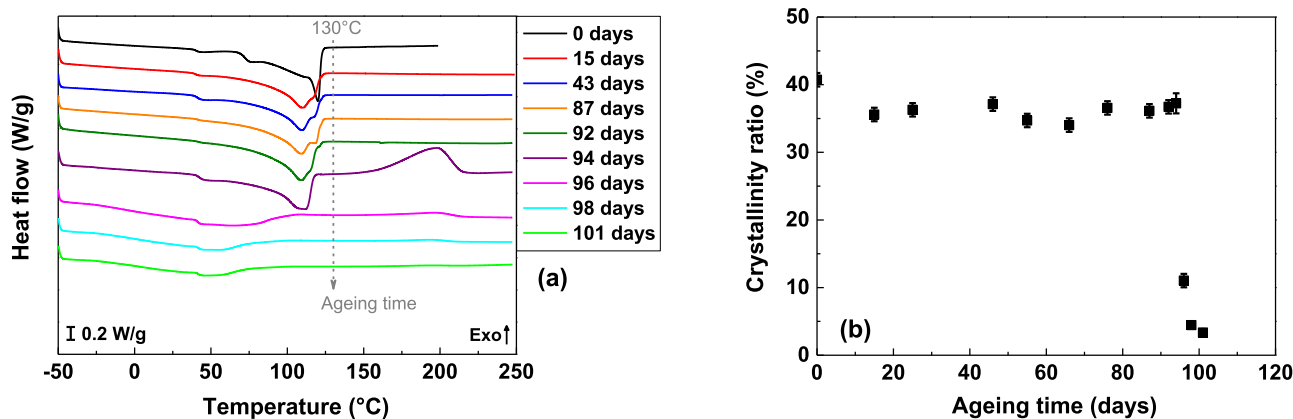


Fig. 7. (a) Changes in the DSC thermogram (heating rate of $10^{\circ}\text{C}\cdot\text{min}^{-1}$) of the Si-XLPE stabilised with 0.6 wt% of Irganox 1076® during its thermal ageing at 130°C in air. (b) Corresponding changes in the crystallinity ratio of the polymer with time of exposure.

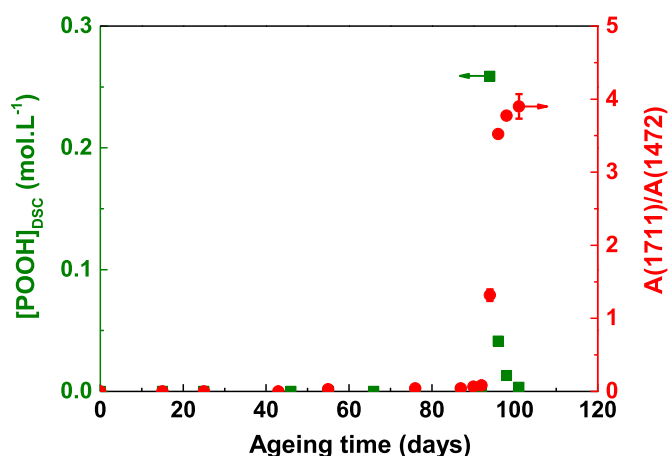


Fig. 8. Comparison of the changes in (■) the hydroperoxide concentration (determined by DSC under N_2) and (●) the normalised absorbance of carbonyl species (determined by FTIR spectroscopy in ATR mode at 1711 cm^{-1}) during the thermal ageing at 130°C in air of the Si-XLPE stabilised with 0.6 wt% of Irganox 1076®.

the phenolic AO have been consumed by the stabilisation reactions ($\text{OIT} = 0$), thus leading to an accumulation of POOH into the polymer matrix. It is expected that the AO consumption suddenly auto-accelerates as their concentration vanishes, because these latter have more and more difficulties in blocking the polymer oxidation. It is thus not surprising to suddenly detect POOH without any prior warning.

However, at 130°C , the thermal decomposition of POOH (both by unimolecular and bimolecular pathways) is also very rapid. That is the reason why, as soon as they are detected, POOH already start to disappear from the polymer matrix, thus leading to the formation of radical species, which initiate the oxidation reaction. It should be remembered that the thermal decomposition of POOH is all the more rapid as their concentration is high [6,14,53]. For such a high POOH concentration, around $0.3\text{ mol}\cdot\text{L}^{-1}$, it is thus not surprising not to observe an additional induction period specific to the thermal oxidation of the polymer matrix without antioxidant. On the contrary, the polymer oxidation starts with a very high rate as soon as the phenolic AO have been totally consumed by the stabilisation reactions. Indeed, in Fig. 8, it can be observed an almost vertical increase of the carbonyl build-up. No doubt that, under these conditions, the embrittlement criterion of the polymer matrix will be quickly reached and the oxidation induction time would correspond to the polymer lifetime.

3.2.3. Chemical changes in the bulk

3.2.3.1. FTIR spectroscopy in transmission mode. Finally, the FTIR spectra obtained in transmission mode before and after different durations at 130°C in air are shown in Fig. 9. Both hydroxyls (between 3750 and 3550 cm^{-1}) and carbonyls spectral regions (between 1850 and 1600 cm^{-1}) were normalised with the sample thickness to better visualize the changes.

In the hydroxyls region, three main IR bands can be observed for the unaged material. The IR band at 3605 cm^{-1} is due to two-phonon combinations of CH_2 asymmetric stretching and rocking modes, i.e. to the PE matrix itself [54–56]. In addition, the IR band at 3695 cm^{-1} is attributed to the O-H stretching of residual silanol functions (Si-OH) [48,57–59]. Indeed, crosslinking via the silane route first occurs by the hydrolysis of alkoxy silane functions (Si-OR, where R is an alkyl function) into silanol functions, which then react with one another through a condensation reaction to form the siloxane functions (Si-O-Si). Finally, the third IR band at 3649 cm^{-1} corresponds to the phenol functions of Irganox 1076®.

In the carbonyls region, only one IR band is initially detected at 1741 cm^{-1} , corresponding to the ester function of AO. As expected, the IR positions of the phenol and ester functions of Irganox 1076® corresponds well to those of dissolved AO (in dispersed state), as recalled in Table 1.

As already observed for the pure Irganox 1076® powder, IR bands around 1633 cm^{-1} and 1720 cm^{-1} appear during thermal ageing. As previously mentioned, these IR bands can respectively be assigned to quinone methides and hydroxycinnamates, both resulting from the chemical consumption of AO. Indeed, as the ageing proceeds, it can be observed that the IR band of the phenol functions decreases continuously. Meanwhile, in the carbonyls region, a broad IR band between 1730 and 1700 cm^{-1} (several types of carbonyl species) appears and increases with ageing time. This composite band is mostly visible from 43 days of exposure. Then, after 94 days of exposure, another large IR band between 3000 and 3500 cm^{-1} (O-H bonds of hydroperoxides, alcohols and/or carboxylic acids) is in turn observed, and the FTIR spectra saturate in the carbonyls region due to a significant formation of carbonyl species, i.e. high conversion ratios of the polymer oxidation reaction, as already suggested by our previous results.

Using the Beer-Lambert's law (Eq. 1), the concentrations of phenols, esters and cinnamates were estimated from FTIR results. The molar extinction coefficients of the phenol and ester functions of the AO dissolved in the PE matrix were already determined in a previous work [29] and are about $100\text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ and $490\text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$, respectively. In contrast, the molar extinction coefficients of quinones and cinnamates were selected from liter-

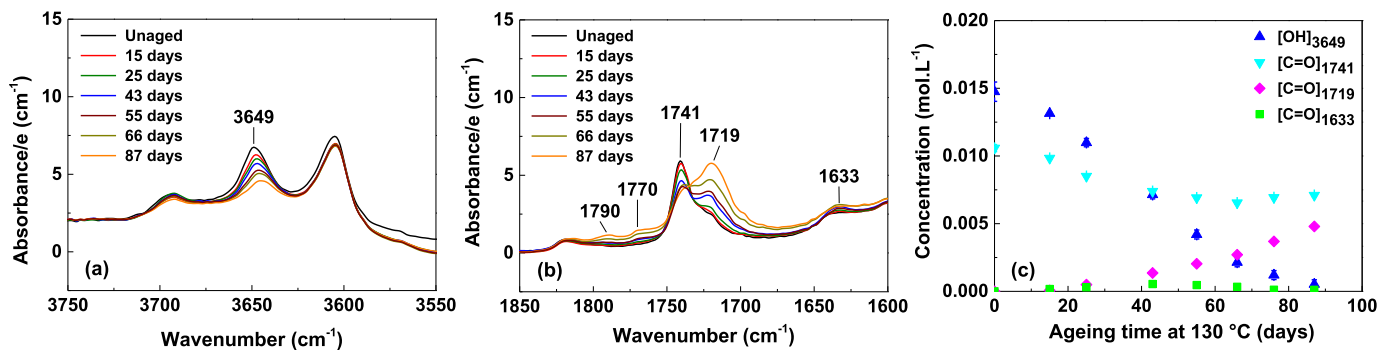


Fig. 9. Changes in the two normalised regions of interest of the FTIR spectrum in transmission mode of the Si-XLPE stabilised with 0.6 wt% of Irganox 1076® during its thermal ageing at 130°C in air: (a) between 3750-3550 cm^{-1} , and (b) between 1850-1600 cm^{-1} . (c) Corresponding changes in the concentrations of the characteristic functions of AO and its degradation products with time of exposure: (\blacktriangle) phenols, (\blacktriangledown) esters, (\blacklozenge) cinnamates and (\blacksquare) quinones.

ature. According to Flett [51], aromatic ketones and esters have molar extinction coefficients typically ranged between 280-490 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ and 500-700 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$, respectively. In a first approximation, orders of magnitude of 400 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ and 700 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ were respectively considered in this study.

The concentrations of these four chemical functions were plotted against time of exposure in Fig. 9c. Three kinetic regimes can clearly be distinguished:

- i) In the first regime (until about 43 days of exposure), the concentrations of both phenol and ester functions of AO decrease with ageing time, which can be due to its chemical consumption and/or its physical loss through evaporation. In this same regime, some quinone methides (IR band at 1633 cm^{-1}) and cinnamates (IR band at 1719 cm^{-1}) are also detected, indicating that the chemical consumption of AO is really occurring. It should be noted that the saturated ester function (initially present in Irganox 1076®) is kept in quinone methides, but has disappeared in cinnamates which, on the contrary, contain an unsaturated ester. Hence, a decrease in the saturated ester functions (IR band at 1741 cm^{-1}) is either due to the AO physical loss through evaporation and/or the formation of cinnamates. At 25 days of exposure, for instance, the concentrations of formed quinone methides and cinnamates are estimated to be around $3 \times 10^{-4} \text{ mol}\cdot\text{L}^{-1}$ and $5 \times 10^{-4} \text{ mol}\cdot\text{L}^{-1}$, respectively. Meanwhile, the decreases in the concentrations of phenol and ester functions of the AO are estimated at approximately $4 \times 10^{-3} \text{ mol}\cdot\text{L}^{-1}$ and $2 \times 10^{-3} \text{ mol}\cdot\text{L}^{-1}$, respectively. These results hence suggest that, in this first regime, the evaporation of AO predominates largely over its chemical consumption since its characteristic functions (phenol and ester) disappear one order of magnitude faster than the formation of its degradation products (quinone methides and cinnamates). In fact, let us remind that in stabilised polymers, the chemical consumption of phenolic AO generally is due to the reaction of AO with peroxy radicals (POO^\bullet) resulting from the polymer oxidation. However, in stabilised polymers, the initial concentration of POOH (and thus, of peroxy radicals) is generally very low, which could explain the few chemical consumption of AO observed in this first regime. Moreover, as quinone methides and cinnamates were also detected during the thermal ageing at 130°C in air of the pure AO powder, the chemical consumption of AO, from the direct reaction of AO with molecular oxygen (which is a bi-radical), can also not be totally excluded; although such a reaction is generally considered negligible in the case of hydrocarbon poly-

mers exposed to such conditions (i.e. typically at $T < 200^\circ\text{C}$ and under low O_2 partial pressures) [26,42,60].

- ii) In the second regime (between 43 and 66 days of exposure), the concentration of esters (IR band at 1741 cm^{-1}) seems to remain relatively constant around $7 \times 10^{-3} \text{ mol}\cdot\text{L}^{-1}$ although the concentration of phenols continues to decrease with ageing time. In the same time, the concentration of quinone methides decreases, while the concentration of cinnamates (IR band at 1719 cm^{-1}) increases more significantly. Let us recall that, according to literature, and reminded in Scheme 1, cinnamates are formed by an intramolecular rearrangement of quinone methides, which could explain the opposite evolutions in the concentrations of these two chemical species. In addition, it is also noteworthy that, according to literature [61], alkylperoxycyclohexadienone species would also be formed in high yield from the reaction between phenolic AO and peroxy radicals. The carbonyl moiety of these species can, according to literature, typically be detected around 1695-1712 cm^{-1} [61]. The apparent increase of cinnamates concentration (IR band at 1719 cm^{-1}) could, hence, also be due to the formation of alkylperoxycyclohexadienones, which absorb in almost the same spectral range. Indeed, it can be observed that the corresponding carbonyl IR massif is relatively broad and thus, might contain the contributions of several species. In addition, let us note that saturated ester functions are present in alkylperoxycyclohexadienones, but not in cinnamates. Thus, the formation of alkylperoxycyclohexadienones does not involve the depletion of saturated ester functions (IR band at 1741 cm^{-1}), unlike the formation of cinnamates. As the increase in the IR band at 1719 cm^{-1} is accompanied with the depletion of phenol functions but no apparent change in the concentration of saturated esters, the formation of alkylperoxycyclohexadienones is also suspected to occur. Moreover, as reminded in Fig. 1 and Scheme 1, the formation of alkylperoxycyclohexadienones requires the presence of peroxy radicals, which are formed from the polymer oxidation. According to literature [26,62], the chemical consumption of phenolic AO by peroxy radicals should be observed after an induction period, as the initial concentration of POOH (and thus, peroxy radicals POO^\bullet) is generally very low, if not near zero. Finally, it should be mentioned that, in the long term, the IR bands of saturated esters (at 1741 cm^{-1}) and cinnamates (at 1719 cm^{-1}) should also be impacted by the formation of some oxidation products from the polymer matrix, such as aldehydes and ketones, which respectively absorb around 1730-1740 cm^{-1} and 1718-1720 cm^{-1} [37,63,64].

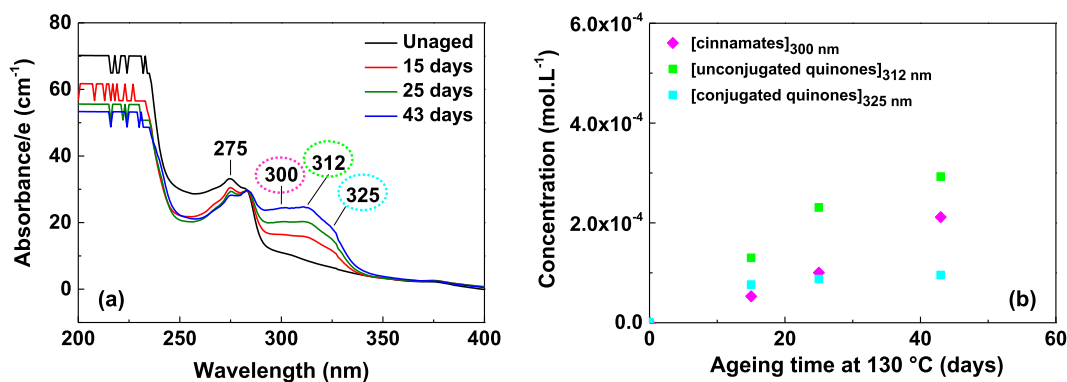


Fig. 10. (a) Changes in the normalised UV spectrum of the Si-XLPE stabilised with 0.6 wt% of Irganox 1076® during its the thermal ageing at 130°C in air; (b) Corresponding changes in the concentrations of (♦) cinnamates, (■) conjugated quinones and (■) unconjugated quinones with time of exposure.

iii) In the last regime (i.e. beyond 66 days of exposure), the IR bands around 1770-1790 cm^{-1} slightly increase, presumably due to the formation of lactones from the polymer oxidation. A similar behaviour was observed in a previous work on the thermal ageing of a Si-XLPE stabilised with a thiodipropionate AO [37].

3.2.3.2. UV spectroscopy in transmission mode. As shown previously, quinones can be detected around 1633 cm^{-1} by FTIR spectroscopy. However, as shown in Fig. 5b and Fig. 9b, their absorbance is generally very low. In literature, UV spectroscopy is hence more often used to detect them as their molar extinction coefficient in UV is much higher. In fact, Klemchuk and Horng [36] found molar extinction coefficients of about 35 000 (at 314 nm) and 45 000 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ (at 322 nm) for unconjugated and conjugated bisquinonemethides, respectively.

The UV spectra obtained before and after different durations at 130°C in air are shown in Fig. 10a. These spectra were also normalised with the film thickness (e) to better visualize the changes. In UV spectroscopy, in addition to quinone methides, the aromatic phenols and hydroxycinnamates can also be detected around 275 nm and 300 nm [36], respectively.

As for FTIR spectroscopy, the Beer-Lambert's law (Eq. 1) was used to quantify the chemical species detected by UV spectroscopy. In particular, the concentration of quinone methides and cinnamates were estimated using the following molar extinction coefficients from literature: 35 000 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ for conjugated quinones (UV band at 325 nm), 45 000 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ for unconjugated quinones (UV band at 312 nm) [36], and 35 000 $\text{L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ for hydroxycinnamates (UV band at 300 nm) [36]. The calculated concentrations were then plotted against time of exposure in Fig. 10b.

The results obtained by UV spectroscopy suggest that quinone methides are the main degradation products of AO, with a total concentration of about 3×10^{-4} $\text{mol}\cdot\text{L}^{-1}$ after 25 days of exposure at 130°C, while the concentration of cinnamates is about 1×10^{-4} $\text{mol}\cdot\text{L}^{-1}$. This is the same concentration of quinones previously determined by FTIR spectroscopy. In contrast, the concentration of cinnamates is about five times lower than previously determined by FTIR spectroscopy. This concentration difference might be due to the additional contribution of carbonyl species due to polymer oxidation in FTIR spectroscopy, as previously mentioned. However, it should be pointed out that this difference is totally negligible at low conversion ratios, for instance after 15 days of exposure at 130°C. In this case, indeed, the concentrations of quinones and cinnamates estimated by FTIR and UV spectroscopies are of the same order of magnitude, thus giving a good confidence in the values of the molar extinction coefficients chosen for this study.

Unfortunately, for heterogeneously stabilised samples (i.e. with a concentration of Irganox 1076® higher than 0.9 wt%), UV spectroscopy could not be used to quantify the AO and its degradation products due to the saturation of the corresponding UV absorption bands. Thus, for these samples, only FTIR spectroscopy was used.

3.3. Physico-chemical characterization of the thermal ageing of the Si-XLPE stabilised with an excess of Irganox 1076®

The thermal ageing of a Si-XLPE stabilised with 2.1 wt% of Irganox 1076® was investigated in air at three different temperatures: 87°C, 110°C, and 130°C, using the same general methodology detailed in Fig. 4.

3.3.1. Surface analysis (FTIR spectroscopy in ATR mode vs. Optical microscopy)

The FTIR spectra in ATR mode obtained before and after different durations at 87°C, 110°C, and 130°C in air are shown in Fig. 11, Fig. 12, and Fig. 13, respectively. As for the homogeneously stabilised sample, a special attention was paid to the three spectral regions of interest normalised with the IR band at 1472 cm^{-1} .

AO crystals are initially detected on the film surface. Indeed, the initial FTIR spectrum contains IR bands at 3639 and 1733 cm^{-1} corresponding to the phenol and ester functions of AO under the crystalline Form I, respectively. As a reminder, the material under study was stabilised by an initial concentration of Irganox 1076® above its solubility threshold (about 0.9 wt%) in PE-based materials [29,33]. These AO in excess exuded and recrystallized at the sample surface during the processing operation [33].

As the thermal ageing proceeds, the IR bands of AO crystals under the Form I decrease for the three exposure temperatures under investigation, suggesting a decrease of the amount of AO crystals present at the sample surface. In addition, in the case of thermal ageing at 130°C (Fig. 13), the formation of IR bands at 3612 cm^{-1} and 1723 cm^{-1} is observed after 71 days of exposure. These IR bands are characteristic of AO under the crystalline Form III, as recalled in Table 1. Only in this case, an additional IR band is detected at 1288 cm^{-1} , which is not present, neither for the unaged, nor for all other aged samples. According to Saunier et al. [31], this IR band at 1288 cm^{-1} is specific to the crystalline Forms II and III (C-O stretching of the ester) as it is not observed for the crystalline Form I. In order to confirm the presence of this new crystalline form of AO, the sample surface was examined by optical microscopy before and after 71 days of exposure at 130°C. Indeed, it was shown in literature that, depending on the crystalline form involved, Irganox 1076® crystals can have specific shapes and, in particular, the crystalline Form III of Irganox 1076® was observed as needle-shaped crystals [29,31]. On Fig. 14, it can be observed

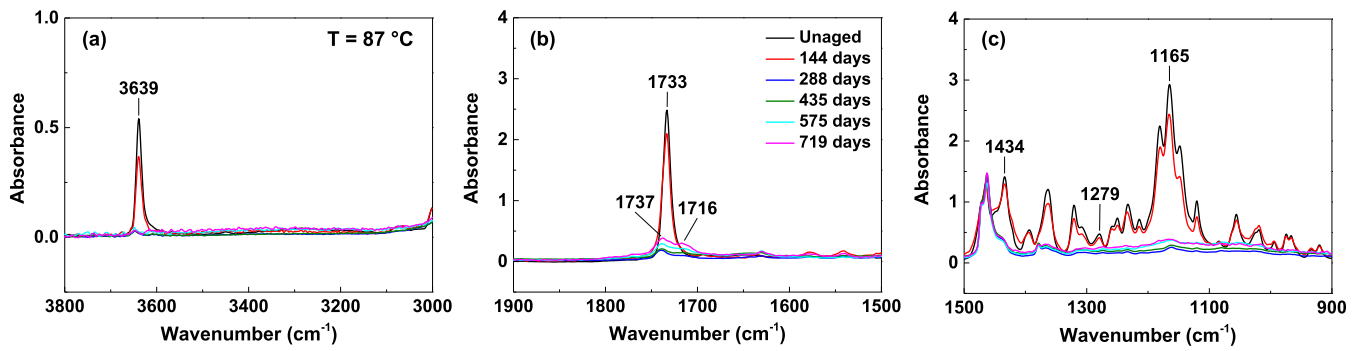


Fig. 11. Changes in the three normalised regions of interest of the FTIR spectrum in ATR mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing at 87°C in air: (a) between 3800-3000 cm^{-1} , (b) between 1900-1500 cm^{-1} , and (c) between 1500-900 cm^{-1} .

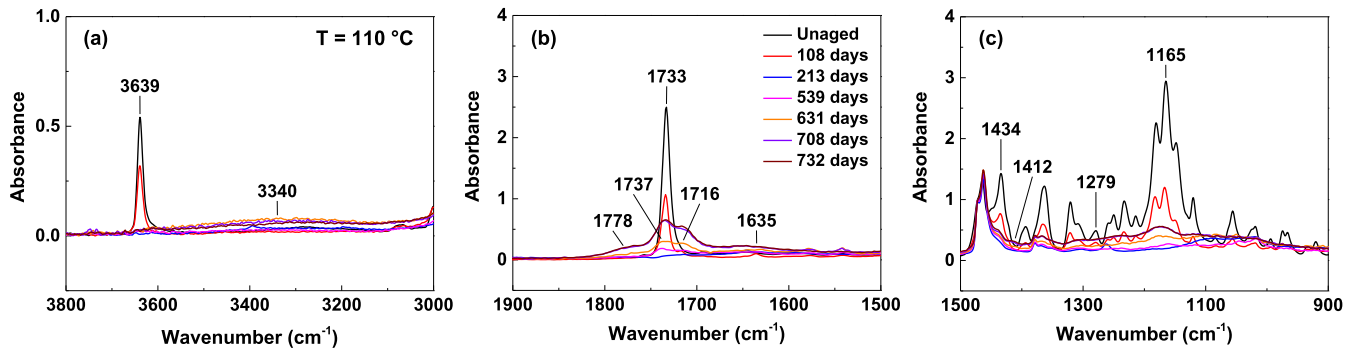


Fig. 12. Changes in the three normalised regions of interest of the FTIR spectrum in ATR mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing at 110°C in air: (a) between 3800-3000 cm^{-1} , (b) between 1900-1500 cm^{-1} , and (c) between 1500-900 cm^{-1} .

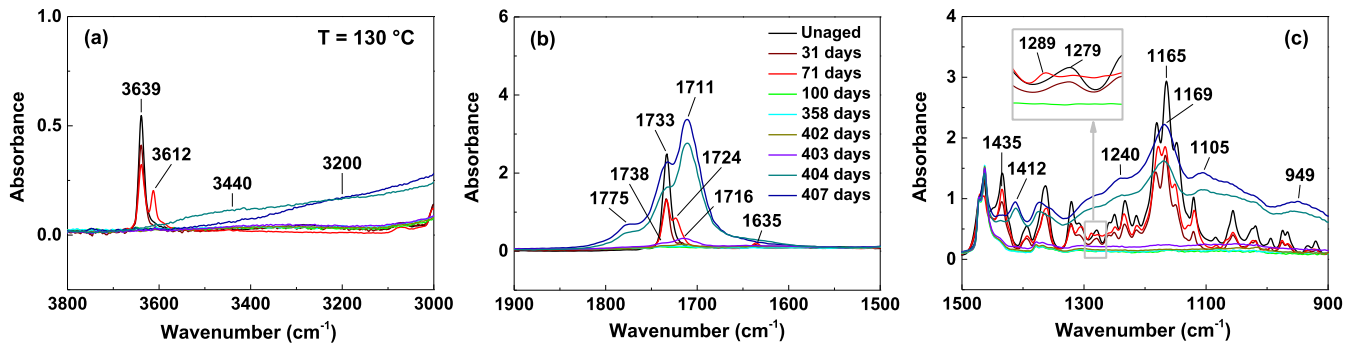


Fig. 13. Changes in the three normalised regions of interest of the FTIR spectrum in ATR mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing at 130°C in air: (a) between 3800-3000 cm^{-1} , (b) between 1900-1500 cm^{-1} , and (c) between 1500-900 cm^{-1} .

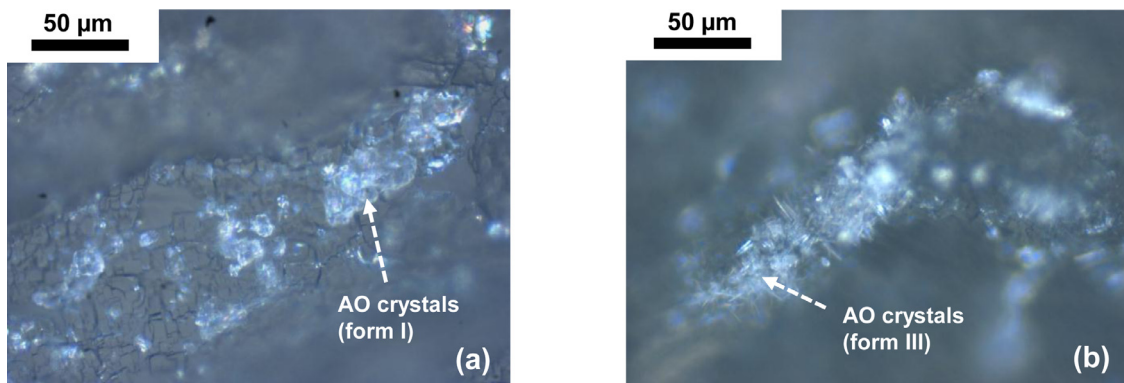


Fig. 14. Micrographs of the surface of the Si-XLPE film stabilised with 2.1 wt% of Irganox 1076®: (a) before and (b) after 71 days of exposure at 130°C under air (with a magnification of $\times 500$).

that before thermal ageing, AO crystals of various sizes (from few microns to a dozen of microns) were present at the sample surface, while after 71 days of exposure at 130°C, needle-shaped crystals are now covering the sample surface. It can be thus concluded that both FTIR spectroscopy in ATR mode and optical microscopy confirm the polymorphism of Irganox 1076® after 71 days of exposure at 130°C in air.

For the three exposure temperatures under investigation and, in particular, for thermal ageing at 130°C, the appearance of an IR band around 1630-1635 cm^{-1} was also noticed during the early periods of exposure. It should be recalled that this IR band was already observed during the thermal ageing at 130°C in air of both of the pure AO powder and the homogeneously stabilised Si-XLPE film. It is assigned to quinone methides resulting from the reaction of the phenol function of AO with peroxy radicals and/or molecular oxygen. These observations hence suggest that, while the amount of AO crystals decreases with time of exposure on the sample surface, due to a further dissolution of AO into the polymer matrix favoured by the sample exposure to higher temperatures (the melting point of AO crystals is about 50°C) and/or its evaporation, a non-negligible part of the AO is also chemically consumed. Unfortunately, at this stage of the investigations, it is not possible to know in which physical state (i.e. dissolved or crystallised) are the AO the most prone to chemical consumption. This is surely a difficult problem to solve because it cannot be excluded that a part of the degradation products (e.g. quinone methides and cinnamates) formed in the AO crystals dissolves into the polymer matrix during thermal ageing, since these latter are less polar than the initial phenolic AO.

Then, as the thermal ageing proceeds, no IR band is detected for AO at the sample surface, in contrast, a slight growth in the IR bands in the carbonyls spectral range (around 1737 cm^{-1} and 1716 cm^{-1}) is noticed. These latter are typically attributed to aldehydes and ketones from PE oxidation. The increase in their absorbance with time of exposure is clearly a sign of the onset of polymer oxidation at the sample surface, as these IR bands were also observed during the thermal ageing of the pure Si-XLPE. Indeed, as for a linear PE [65,66], the thermal oxidation of Si-XLPE leads to the formation of a wide variety of carbonyl species such as: carboxylic acids and saturated ketones (typically between 1700 and 1720 cm^{-1}), aldehydes and esters (between 1730 and 1740 cm^{-1}), and γ -lactones (around 1770-1790 cm^{-1}) [48,65-67].

For thermal ageing at 130°C in air, in particular, these IR bands first slowly increase up to 404 days of exposure for which, as previously observed for the homogeneously stabilised sample, oxidation accelerates suddenly. In contrast, for the other two temperatures of exposure (87 and 110°C), although some oxidation products of polymer are detected by FTIR spectroscopy in ATR mode, their absorbance remain quite low. It should be noted that the OIT values measured at 205°C for most samples aged at 87°C and 110°C are still non-null: about 100 min after 720 days of exposure at 87°C and about 20 min after 740 days at 110°C. They correspond to AO concentrations of about $3 \times 10^{-2} \text{ mol.L}^{-1}$ and $5 \times 10^{-3} \text{ mol.L}^{-1}$, respectively. These concentrations were estimated using Eq. 2 and the calibration curve obtained for OIT at 205°C in our previous work [33].

3.3.2. Morphological changes (DSC analysis)

Fig. 15 shows the DSC thermograms (under N_2) obtained during the thermal ageing at 87°C, 110°C, and 130°C in air of the heterogeneously stabilised Si-XLPE film.

A modification of the general shape of the melting peak of Si-XLPE can be noticed in the early periods of exposure for the three temperatures under investigation. The same behaviour was observed during the thermal ageing of a Si-XLPE stabilised with a thiopropionate AO [37], and for the Si-XLPE film stabilised with 0.6

wt% of Irganox 1076® aged at 130°C. These modifications are due to the partial or total melting of polymer crystals during thermal ageing. When the temperature of exposure is located within the melting range of Si-XLPE (i.e. at 87°C and 110°C), almost all the smallest crystals having a melting point lower than the temperature of exposure are transformed into larger crystals which are stable at the temperature of exposure, presumably by coalescence [68,69]. As a consequence, the lower part of the melting peak disappears, the melting peak becomes narrower and can be shifted towards higher temperatures (in particular at 110°C). Under these conditions, it is not surprising that almost no recrystallization is observed during the sample cooling to room temperature before DSC analysis, except perhaps at 110°C where a small secondary melting peak can be seen around 90°C on the DSC thermograms.

When the temperature of exposure is higher than the melting range of Si-XLPE (i.e. at 130°C), in contrast, all the polymer crystals are melted during the thermal ageing and a total recrystallization occurs during the sample cooling at room temperature before the DSC analysis. It should be pointed out that the Si-XLPE was crosslinked in its solid state (at 65°C). Therefore, this crosslinked polymer is initially characterized by the same crystalline morphology as the starting linear low density polyethylene (Si-g-LLDPE). However, in molten state, its macromolecular chains are now crosslinked and thus, are characterized by a reduced molecular mobility which hinders the recrystallization of the largest crystals. It is thus not surprising to observe a widening and a shift of the melting peak towards the lower temperatures after recrystallization. These changes in crystalline morphology seem to be accelerated with the polymer oxidation, in particular from 404 days of exposure in air at 130°C.

As the thermal ageing proceeds, no additional modification is observed on the DSC thermograms for the two lowest temperatures, which is consistent with the previous results. Indeed, it should be remembered that, at 87 and 110°C, OIT never becomes null and very few polymer oxidation products are detected by FTIR spectroscopy in ATR mode. In contrast, at 130°C, as already observed for the homogeneously stabilised Si-XLPE (Fig. 7), a broad exothermic peak starts to appear on the DSC thermogram between 130 and 220°C after 402 days of exposure, due to the thermal decomposition of POOH species. The concentration of POOH accumulated in the polymer matrix was estimated using Eq. 5 and plotted against time of exposure in Fig. 16. As for the homogeneously stabilised sample, the normalised absorbance of carbonyl species due to polymer oxidation (determined by FTIR spectroscopy in ATR mode at 1711 cm^{-1}) has also been plotted in the same graph.

As for the homogeneously stabilised sample, as soon as the phenolic AO are completely depleted (OIT = 0), the hydroperoxides are suddenly detected by DSC with a maximum concentration of about 0.1 mol.L^{-1} . These latter immediately initiate the thermal oxidation of the polymer matrix with a very high rate. However, as expected, the induction time of carbonyl species is much longer for the heterogeneously than the homogeneously stabilised sample: 404 days against 94 days of exposure at 130°C in air.

Now, if looking at the crystallinity ratio of Si-XLPE determined with Eq. 4, no significant change is observed during the thermal ageing at 87 and 110°C in air, as shown in Fig. 17. Indeed, there is still no significant sign of polymer oxidation, although some IR bands of oxidation products were detected by FTIR spectroscopy in ATR mode (Fig. 11 and Fig. 12). At 130°C in air, in contrast, an initial slight decrease in the crystallinity ratio (typically from $41 \pm 1 \%$ up to $37 \pm 1 \%$ after 30 days of exposure) is observed presumably due to the polymer post-crosslinking in molten state. This assumption was successfully checked by swelling test in refluxing p-xylene (typically from $71 \pm 1 \%$ up to $76 \pm 1 \%$ after 71 days of exposure). After 404 days of exposure, a sudden catastrophic decrease in the crystallinity ratio is detected due to polymer ox-

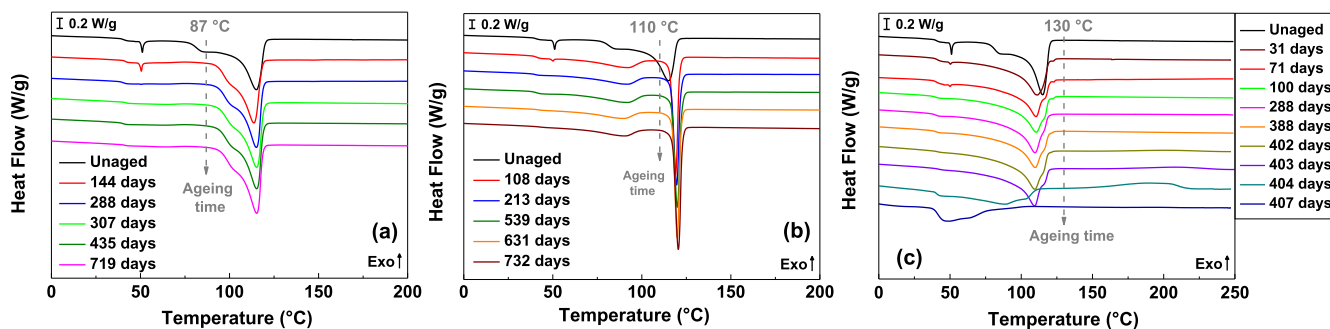


Fig. 15. Changes in the DSC thermogram (heating rate of $10^{\circ}\text{C}\cdot\text{min}^{-1}$) of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing in air at: (a) 87°C , (b) 110°C , and (c) 130°C .

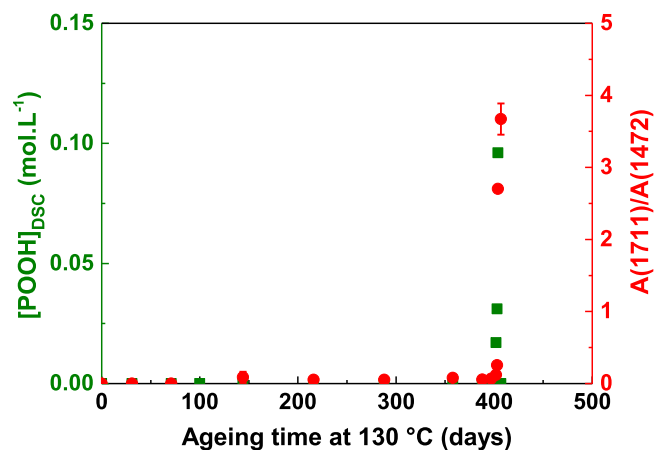


Fig. 16. Comparison of the changes in (■) the hydroperoxide concentration (obtained by DSC under N_2) and (●) the normalised absorbance of carbonyl species (determined by FTIR spectroscopy ATR mode at 1711 cm^{-1}) during the thermal ageing at 130°C in air of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076®.

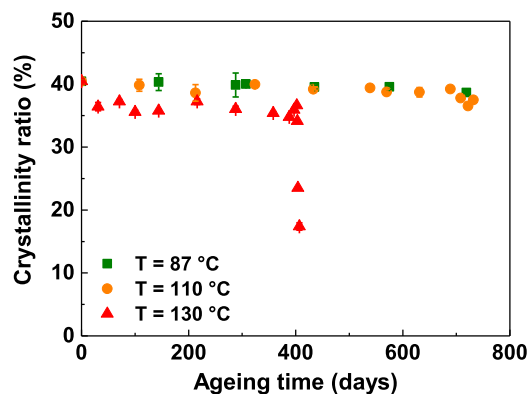


Fig. 17. Changes in the crystallinity ratio of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing in air at (■) 87°C , (●) 110°C , and (▲) 130°C .

ication. Indeed, this duration exactly corresponds to the induction time of carbonyl species detected by FTIR spectroscopy in ATR mode (Fig. 16). It should be remembered that the same behaviour was observed for the homogeneously stabilised Si-XLPE with 0.6 wt% of AO (Fig. 7) and for the pure Si-XLPE both aged at 130°C in air.

Moreover, the DSC thermograms, determined in the first periods of exposure at the three temperatures under investigation in Fig. 15, have been zoomed between 30 and 70°C in Fig. 18. In this temperature range, an endothermic peak is clearly detected at 51°C

on the initial DSC thermogram. It should be remembered that this latter is assigned to the melting of Irganox 1076® crystals under the Form I (Table 2). For the three temperatures of exposure, this peak disappears with ageing time. There is thus a continuous decrease in the concentration of AO crystals during thermal ageing, as already shown by FTIR spectroscopy in ATR mode. When AO crystals are no longer detected at the sample surface by FTIR spectroscopy in ATR mode (Fig. 11, Fig. 12, Fig. 13), there is no longer melting peak of AO crystals on the DSC thermograms. This result is not surprising because, in our previous study [33], it was shown that AO crystals are only present at the sample surface.

Table 3 compiles the concentrations of AO crystals estimated by DSC (using Eq. 3) during the thermal ageing in air at 87°C , 110°C , and 130°C , for the heterogeneously stabilised Si-XLPE film.

Finally, after 71 days of exposure at 130°C , a second endothermic peak is detected by DSC at a slightly lower temperature (47°C) than the melting peak of the initial crystalline Form I of AO (51°C). According to Table 2, this new peak is assigned to the melting of Irganox 1076® crystals under the Form III. This result is also in perfect agreement with the previous FTIR analyses in ATR mode and the previous examinations by optical microscopy which both showed the presence of Irganox 1076® crystals under the Form III at the sample surface for the same thermal exposure conditions.

3.3.3. Chemical changes in the bulk (FTIR spectroscopy in transmission mode)

Finally, the FTIR spectra obtained in transmission mode before and after different durations in air at the three different temperatures under investigation are shown in Fig. 19. Both hydroxyls (between 3750 and 3850 cm^{-1}) and carbonyls spectral regions (between 1850 and 1600 cm^{-1}) were normalised with the sample thickness to better visualize the changes.

In the hydroxyls region, three main IR bands were observed for the unaged material. As a reminder, the IR band at 3605 cm^{-1} is due to the PE matrix itself, while the IR band at 3695 cm^{-1} is due to the O-H stretching from residual silanol functions. The third IR band centred at 3642 cm^{-1} is attributed to the phenol function of Irganox 1076®. In particular, as reminded in Table 1, this latter corresponds to the crystalline Form I of Irganox 1076®.

During the first periods of exposure, the same behaviour is observed for the three temperatures under investigation. At first, it can be noticed that the IR band at 3695 cm^{-1} (Si-OH) decreases due to the post-crosslinking of the polymer matrix via the condensation of residual Si-OH functions. Then, as also previously observed for the thermal ageing of the pure AO powder and the homogeneously stabilised Si-XLPE in air at 130°C , two new IR bands appear between 1650 - 1600 cm^{-1} and around 1720 cm^{-1} on the FTIR spectrum. As a reminder, these two IR bands were respectively attributed to quinone methides and hydroxycinnamates, both

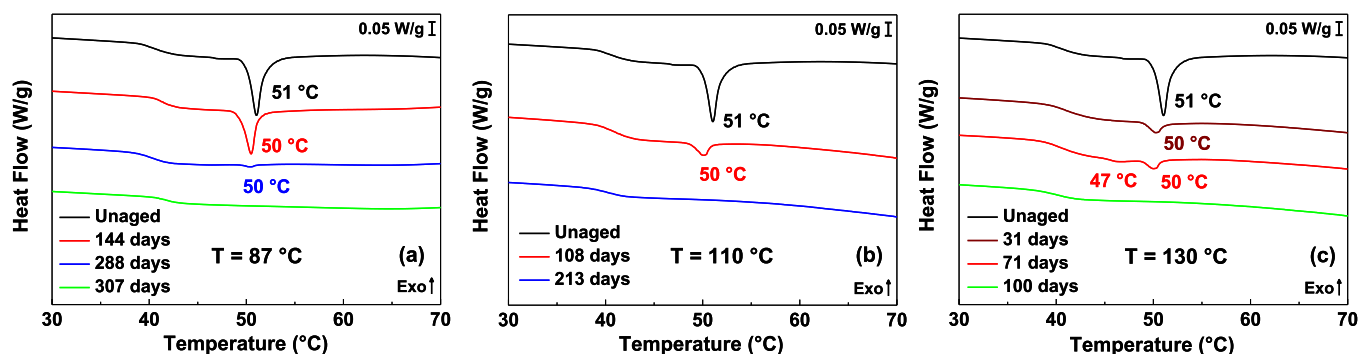


Fig. 18. Zoom between 30 and 70°C of the DSC thermograms determined in the first periods of exposure during the thermal ageing in air at: (a) 87°C, (b) 110°C, and (c) 130°C in Fig. 15.

Table 3

Concentration of AO crystals $[AO]_{insol}$ estimated by DSC for the heterogeneously stabilised Si-XLPE film during its thermal ageing under air at 87°C, 110°C, and 130°C.

T = 87°C		T = 110°C		T = 130°C	
Ageing time (days)	$[AO]_{insol}(\text{mol.L}^{-1})$	Ageing time (days)	$[AO]_{insol}(\text{mol.L}^{-1})$	Ageing time (days)	$[AO]_{insol}(\text{mol.L}^{-1})$
0	$(2.4 \pm 0.4) \times 10^{-2}$	0	$(2.4 \pm 0.4) \times 10^{-2}$	0	$(2.4 \pm 0.4) \times 10^{-2}$
144	$(2.0 \pm 0.2) \times 10^{-2}$	108	$(6.1 \pm 0.1) \times 10^{-3}$	31	$(6.0 \pm 0.1) \times 10^{-3}$
288	$(7.6 \pm 0.1) \times 10^{-4}$	213	/	71	$(4.5 \pm 0.1) \times 10^{-3}$
307	/	-	-	100	/

resulting from the AO chemical consumption. Another interesting observation is the shift of the initial IR bands of the phenol and ester functions of Irganox 1076® during this first ageing stage. This spectral modification was already observed in a previous study on the same material but after thermal treatment under primary vacuum at 70°C [29]. In fact, the material under study contains 2.1 wt% of Irganox 1076®, in which one part of AO is solubilised into the polymer matrix and the second part of AO is exuded under the form of crystals. As reminded in Table 1, it was shown in this previous study that the position of the IR bands of the phenol and ester functions is different depending on the physical state of AO in the polymer matrix (i.e. solubilised or crystallised), but also depending on its crystalline structure. Here, it appears that the IR bands of the phenol and ester functions are initially centred at 3642 and 1736 cm^{-1} , respectively. They thus correspond to the crystalline Form I. Then, as the thermal ageing proceeds, a shift of these two IR bands towards 3649 cm^{-1} and 1741 cm^{-1} , respectively, can be observed. According to Table 1, these two new positions correspond to AO dissolved in PE (i.e. in dispersed state). Consequently, this shift could be due to a decrease of AO crystals concentration during thermal ageing and/or an increase in the soluble fraction of AO into the polymer matrix. Possible explanations could be a further solubilisation of AO into the polymer matrix favoured by the sample exposure to higher temperatures (the melting point of AO crystals is about 50°C), and/or a faster depletion of AO crystals due to physical loss and chemical consumption.

As the thermal ageing proceeds, it can be seen a continuous decrease in the IR bands of the phenol and ester functions of AO at 3648 cm^{-1} and 1740 cm^{-1} , respectively. In addition, for the two highest temperatures of exposure (110 and 130°C), new IR bands already observed in ATR mode appear in the carbonyls region around 1738 cm^{-1} , 1770 cm^{-1} , and 1790 cm^{-1} . As previously mentioned, these new IR bands are due to polymer oxidation, as they were also observed during the thermal ageing of the pure Si-XLPE, but also during the thermal ageing of the Si-XLPE stabilised with a thiopropionate AO [37]. Finally, after 403 days of exposure at the highest temperature (130°C), the IR bands that have appeared in the carbonyls region begin to saturate, while new IR bands appear at 3550 cm^{-1} and between 3500-3200 cm^{-1} in the

hydroxyls region. According to literature, these latter are due to hydroperoxides species (free and H-bonded OH) [66]. It should be remembered that these hydroperoxide species were also detected by DSC analysis (exothermic peak between 130 and 220°C) in Fig. 15c.

The total concentrations of phenol and ester functions of AO were determined by FTIR spectroscopy in transmission mode using Eq. 1 and the calibration curves reported in our previous publications [29,33]. As for the homogeneously stabilised Si-XLPE, the concentrations of quinone methides (IR band at 1633 cm^{-1}) and cinnamates (IR band at 1720 cm^{-1}) were also estimated. The resulting concentrations have been plotted against time of exposure at the three temperatures under investigation in Fig. 20a, Fig. 20b, and Fig. 20c. As for the homogeneously stabilised Si-XLPE, three kinetic regimes can clearly be distinguished.

- i) In the first regime, there is a decrease of both the phenol and ester functions of AO, presumably due to the physical loss of AO through evaporation and/or its chemical consumption. Meanwhile, an initial increase in the concentration of both quinone methides and cinnamates is observed (after which their concentrations remain relatively constant), thus confirming the chemical consumption of AO. It is also noteworthy that the higher the exposure temperature, the higher the concentration of quinone methides and cinnamates. However, the concentration of quinone methides never exceeds $5 \times 10^{-3} \text{ mol.L}^{-1}$, while the decrease in the concentration of the phenol functions of AO is much higher, typically about $2 \times 10^{-2} \text{ mol.L}^{-1}$ after 100 days of exposure at 130°C. As previously concluded for the homogeneously stabilised sample, in this first regime, although some degradation products of AO resulting from its chemical consumption are detected, AO depletion is essentially due to physical loss by evaporation.
- ii) In the second regime only observed at 110°C and 130°C (never reached at 87°C for the exposure duration under study), the concentration of quinone methides decreases, while the concentration of cinnamates starts to slightly increase. Meanwhile, the concentration of phenols decreases further, but the concentration of saturated ester functions

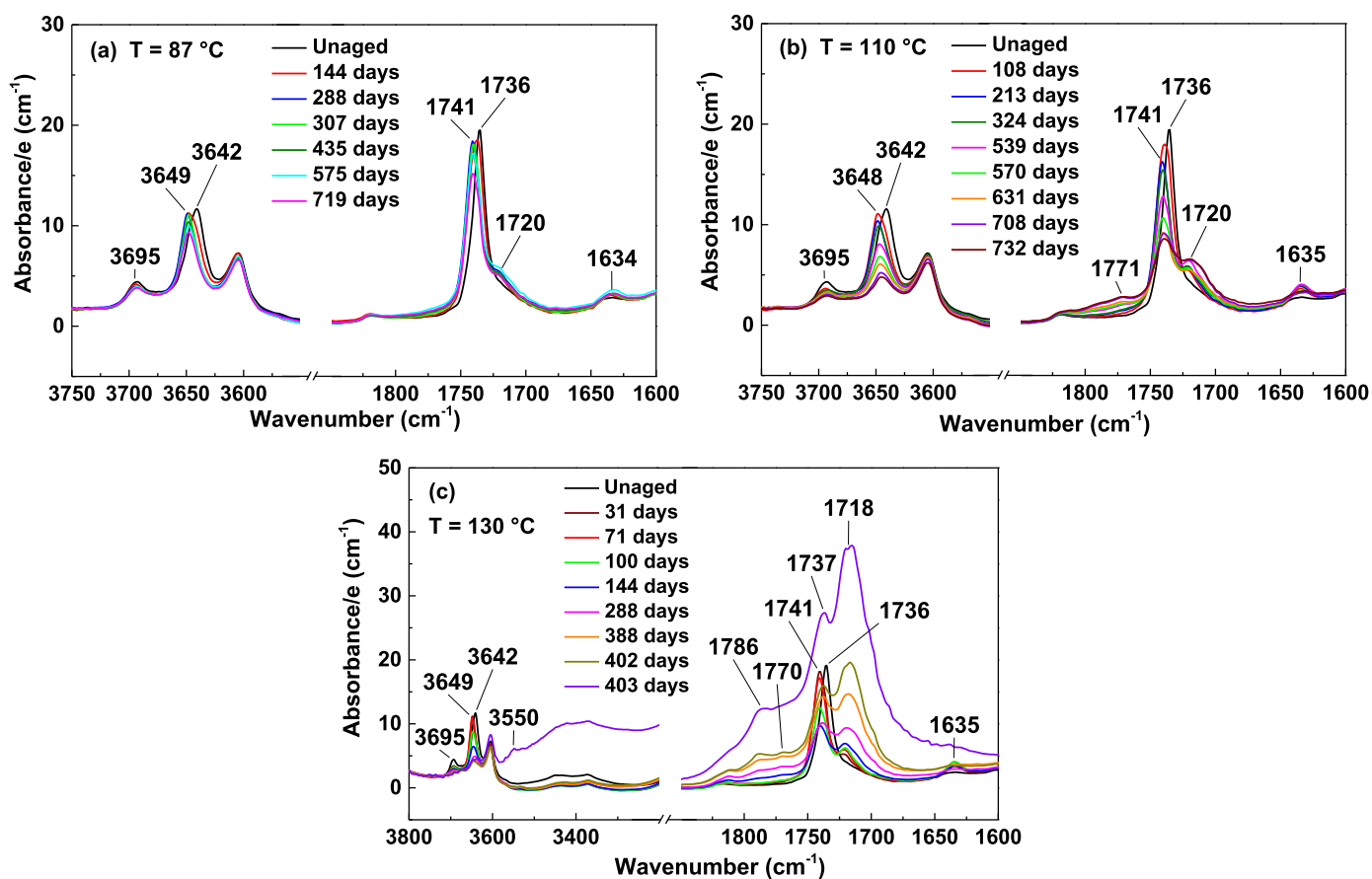


Fig. 19. Changes in the two normalised regions of interest (i.e. between 3850-3550 cm⁻¹ and between 1850-1600 cm⁻¹) of the FTIR spectrum in transmission mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing in air: (a) at 87°C, (b) at 110°C, and (c) at 130°C.

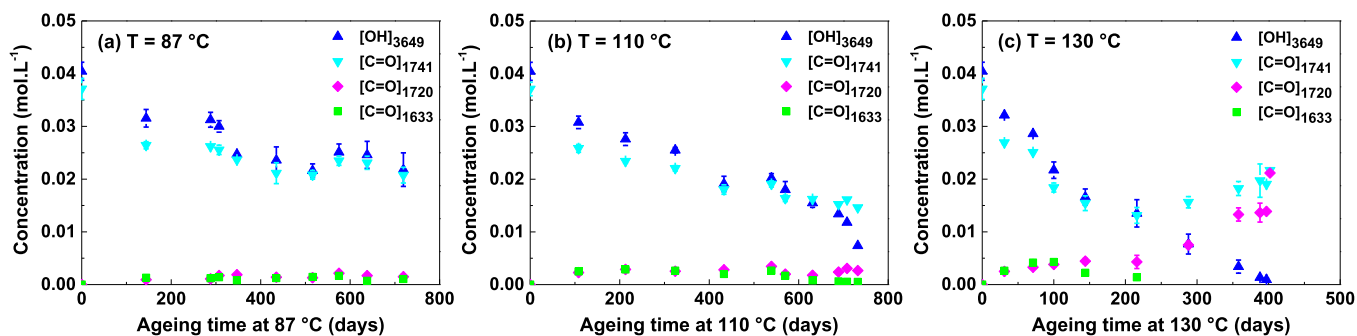


Fig. 20. Changes in the concentrations of the characteristic functions of AO and its degradation products during the thermal ageing in air of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® at: (a) 87°C, (b) 110°C, and (c) 130°C. (▲) phenols, (▼) esters, (◆) cinnamates, and (■) quinones.

(IR band at 1741 cm⁻¹) remains relatively constant. The same behaviour was previously observed for the homogeneously stabilised sample at 130°C, which was presumably due to either the transformation of quinone methides into cinnamates and/or the formation of alkylperoxycyclohexadienones from the reaction between AO and the peroxy radicals resulting from polymer oxidation. Interestingly, for the two highest temperatures of exposure (i.e. at 110°C and 130°C), this second regime seems to be observed when the AO concentration is about 1.3×10^{-2} mol.L⁻¹, which approximately corresponds to a loss of about 70% of the initial AO concentration. It should be mentioned that a second regime for a loss of about 75% of the initial AO concentration was also reported by Tireau et al. [62] for the thermal ageing in air at temperatures ranged between 40

and 80°C for a HDPE stabilised with Irganox 1010 (phenolic AO).

- iii) Then, as the thermal ageing proceeds further at 130°C, a third regime is detected. In this last regime, both the concentration of esters and cinnamates increase while the concentration of phenols is still decreasing with the time of exposure. These apparent increases in the saturated esters (IR band at 1741 cm⁻¹) and cinnamates (IR band around 1720 cm⁻¹) are probably due to the formation of polymer oxidation products, which can overlap the IR bands of the esters and cinnamates as they absorb in the same spectral range. In contrast, the decrease in the concentration of phenols is due to AO depletion. It is also noteworthy that this last regime is observed when the concentration of active AO is relatively low, typically about 3×10^{-3} mol.L⁻¹ after 360 days of expo-

sure at 130°C, which corresponds to a loss of about 90% of the initial AO concentration. The formation of very small IR bands of oxidation products (in particular, ketones at 1718 cm^{-1}) was already detected during the thermal ageing in air at temperatures between 87°C and 130°C of the Si-XLPE stabilised with a thiodipropionate AO. In particular, oxidation products were observed when OIT at 195°C is relatively low: typically about 20 min, which corresponds to a residual concentration of active functions of AO (i.e. sulphides in that case) of about $10^{-3} \text{ mol.L}^{-1}$. However, the detected concentration of ketones was still relatively low (about 10^{-3} - $10^{-2} \text{ mol.L}^{-1}$), thus suggesting the first signs of the polymer oxidation [37]. In addition, for PE stabilised with Irganox 1076®, Moisan and Lever [70] also observed the appearance of oxidation products (IR band of ketones at 1720 cm^{-1}) although AO were still present in the material. Here again, the concentrations of detected oxidation products were very low. These results hence suggest that the polymer oxidation is starting although few AO are still present in the material, which seems to indicate a heterogeneous oxidation of the polymer matrix.

Finally, after 403 days of exposure in air at 130°C, OIT value at 205°C is zero, i.e. no more active AO is present in the Si-XLPE sample. Thenceforth, hydroperoxide species start to be detected (both by FTIR spectroscopy and DSC analysis), and the FTIR spectra (in transmission mode) in the carbonyls region start to saturate. This bands saturation is to be related to the significant increase in the normalised absorbance of carbonyl species measured by FTIR spectroscopy in ATR mode and the significant decrease in the crystallinity ratio of Si-XLPE detected by DSC analysis, both suggesting a catastrophic auto-acceleration of the polymer oxidation. For the two lowest temperatures (i.e. at 87°C and 110°C), however, even after 700 days of exposure, no significant sign of polymer oxidation is still detected.

3.3.4. Analysis after removal of antioxidant crystals at the sample surface

3.3.4.1. *Surface analysis (FTIR spectroscopy in ATR mode).* To check the validity of the assumption, whereby a further solubilisation of AO is occurring during the thermal ageing in air at high temperature, the surface of some aged samples was cleaned using acetone in order to remove the exuded AO crystals. In our previous study [33], it was indeed shown that all AO in excess were exuded at the sample surface and that surface cleaning with acetone is an efficient method to remove all the AO crystals. This treatment will hence enable to characterize only the dissolved AO through the same experimental methodology used until then.

In Fig. 21 are shown the FTIR spectra obtained in ATR mode (i.e. on sample surface) of the heterogeneously stabilised sample aged in air at 130°C, before and after its surface cleaning with acetone. It can be observed that, after surface cleaning, no more IR bands of phenol and ester functions of AO and its degradation products (i.e. quinone methides at 1636 cm^{-1} , and cinnamates at 1718 cm^{-1}) were detected. This results hence suggest that all the AO crystals, and all the degradation products of AO formed at the sample surface, were well removed by cleaning surface with acetone.

3.3.4.2. *Bulk analysis (FTIR spectroscopy in transmission mode).* In Fig. 22, are shown the FTIR spectra obtained in transmission mode (i.e. analysis of the whole thickness) of the same thermally aged samples, before and after surface cleaning with acetone. As expected, after surface cleaning, there is a decrease in the absorbance of the IR bands of AO, due to the removal of AO crystals from the

sample surface, thus leaving only the AO dissolved in the polymer matrix, as already observed in our previous work [33]. Moreover, it can be observed that, after surface cleaning, the IR bands of both the phenol and ester functions of AO are more intense for the aged samples than for the unaged sample, thus suggesting a further dissolution of AO crystals into the polymer matrix during thermal exposure. In fact, at ambient temperature, the sample stabilised with 2.1 wt% of Irganox 1076® was well above the solubility threshold of AO in the polymer matrix (which was found to be about $1.6 \times 10^{-2} \text{ mol.L}^{-1}$ in our previous work [29]), meaning that a high concentration of AO was in excess and under the form of AO crystals exuded at the sample surface. It should be recalled that the AO solubility in a polymer matrix is an increasing function of temperature [27,71,72], which means that the solubility threshold of AO is necessarily higher during thermal exposure at 130°C than at ambient temperature. It should also not be forgotten that, at this temperature, both AO and polymer are melted, which also promotes the further dissolution of AO as these latter are only soluble in the polymer amorphous phase.

In addition, as AO are only soluble in the polymer amorphous phase, an increase in the fraction of the polymer amorphous phase should result in a further dissolution of AO into the polymer matrix. In fact, here, and in particular for thermal ageing at 130°C, a decrease in the polymer crystallinity ratio was observed during the first stage of the thermal exposure, which could thus have induced the slight increase in the dissolved AO concentration. Furthermore, a last explanation for this increase would involve a progressive modification of the global polarity of the samples due to the formation of oxidation products (polar species) from AO and/or polymer during thermal exposure.

Finally, in Fig. 22, it can also be noticed that, after surface cleaning, the absorbance of the IR bands of quinones and cinnamates decreases slightly, but that these two degradation products of AO are still present. It should be recalled that the quinones and cinnamates detected after thermal exposure on the sample surface were all removed by surface cleaning (Fig. 21). Therefore, the observed decrease in these two IR bands is just due to the removal of the degradation products of AO from the sample surface. It can be thus concluded that the remaining quinones and cinnamates are dissolved in the sample bulk.

Table 4 compiles the total concentrations of AO (i.e. both dissolved and crystallised) and the concentrations of AO crystals estimated by FTIR spectroscopy and DSC for the heterogeneously stabilised sample before and after thermal ageing in air at 130°C, and before and after surface cleaning with acetone.

The results are quite consistent with a further solubilisation of AO crystals into the polymer matrix during thermal exposure, as the concentration of dissolved AO in the polymer matrix (i.e. determined by FTIR spectroscopy after surface cleaning) was higher during than before the first stage of thermal exposure. In addition, the differences between the AO concentrations determined before and after surface cleaning (by FTIR spectroscopy) are very close to the concentrations of AO crystals (determined by DSC analysis). It should also be pointed out that, after 100 days in air at 130°C, no more AO crystals were detected by DSC analysis (i.e. only dissolved AO were present), and that a similar concentration of AO was determined (by FTIR spectroscopy) for the sample aged for 71 days in air at 130°C after surface cleaning. During the exposure at high temperature, the AO in excess can hence also participate to the stabilisation process as they can further solubilise into the polymer matrix, ultimately resulting into a quasi-homogeneous distribution of AO.

For the two lower temperatures under investigation (i.e. at 87°C and 110°C), the surface cleaning was not performed. However, similar results can be suspected as the total concentration of AO (determined by FTIR spectroscopy in transmission mode) in aged sam-

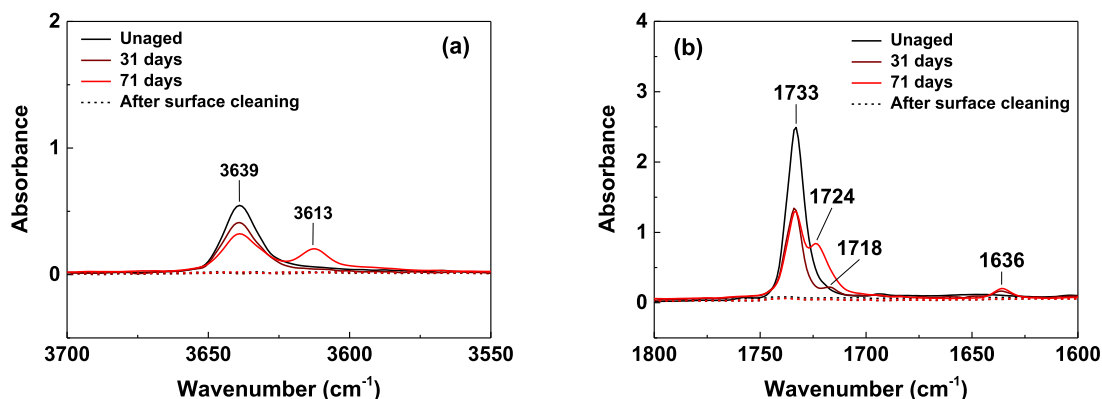


Fig. 21. Changes in the two normalised regions of interest of the FTIR spectrum in ATR mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing at 130°C in air, before (solid lines) and after surface cleaning with acetone (dashed lines): (a) between 3700-3550 cm^{-1} and (b) between 1800-1600 cm^{-1} .

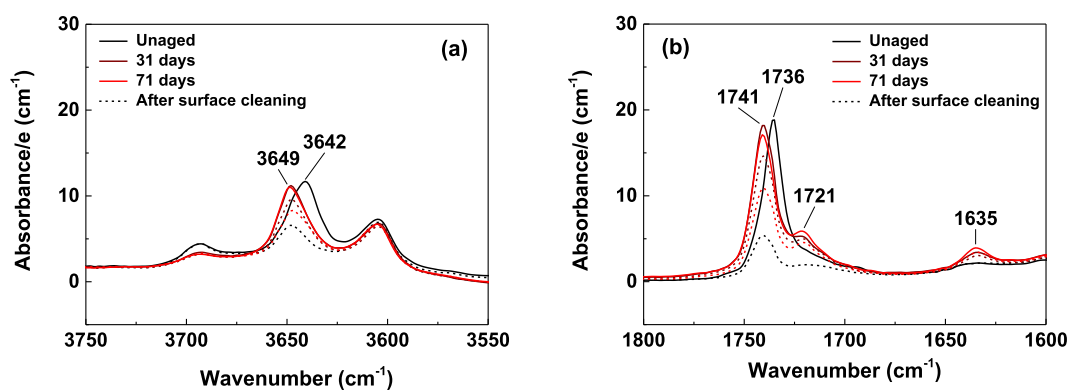


Fig. 22. Changes in the two normalised regions of interest of the FTIR spectrum in transmission mode of the Si-XLPE stabilised with 2.1 wt% of Irganox 1076® during its thermal ageing at 130°C in air, before (solid lines) and after surface cleaning with acetone (dashed lines): (a) between 3750-3550 cm^{-1} and (b) between 1800-1600 cm^{-1} .

Table 4

Concentration of AO estimated by FTIR and DSC for the heterogeneously stabilised Si-XLPE film during its thermal ageing under air at 130°C, before and after surface cleaning with acetone.

Sample		FTIR (Transmission) [AO] _{tot} (mol.L ⁻¹) (from phenol IR band)	DSC (N ₂) [AO] _{insol} (mol.L ⁻¹)
Unaged	No surface cleaning	$(4.1 \pm 0.2) \times 10^{-2}$	$(2.4 \pm 0.4) \times 10^{-2}$
After 31 days at 130°C	After surface cleaning	$(1.6 \pm 0.1) \times 10^{-2}$	/
	No surface cleaning	$(3.2 \pm 0.1) \times 10^{-2}$	$(6.0 \pm 0.1) \times 10^{-3}$
After 71 days at 130°C	After surface cleaning	$(2.6 \pm 0.1) \times 10^{-2}$	/
	No surface cleaning	$(2.9 \pm 0.1) \times 10^{-2}$	$(4.5 \pm 0.1) \times 10^{-3}$
After 100 days at 130°C	After surface cleaning	$(2.1 \pm 0.1) \times 10^{-2}$	/
	No surface cleaning	$(2.2 \pm 0.2) \times 10^{-2}$	/

Table 5

Concentration of AO estimated by FTIR and DSC for the heterogeneously stabilised Si-XLPE film during its thermal ageing under air at 87°C and 110°C.

Sample	FTIR (Transmission) [AO] _{tot} (mol.L ⁻¹) (from phenol IR band)	DSC (N ₂) [AO] _{insol} (mol.L ⁻¹)
After 307 days at 87°C	$(3.0 \pm 0.1) \times 10^{-2}$	/
After 213 days at 110°C	$(2.8 \pm 0.2) \times 10^{-2}$	/

ples, where no more AO crystals are no longer detected (both by DSC and FTIR spectroscopy in ATR mode), is also higher than the concentration of dissolved AO in the unaged sample (1.6×10^{-2} mol.L⁻¹), as shown in Table 5.

4. Conclusion

The thermal ageing in air between 87 and 130°C of a Si-XLPE stabilised with an excess of phenolic antioxidant (Irganox 1076®) was investigated using several complementary techniques. Given the microstructural complexity of the material, which contains both dissolved and crystallised AO, simpler materials were first investigated to separate and elucidate the behaviour of pure AO crystals and dissolved AO in the Si-XLPE matrix.

The thermal ageing in air at 130°C of the pure AO crystals was investigated using the as-received commercial powder. Through FTIR spectroscopy analysis, several chemical changes were detected. Along with the decrease of the AO IR bands, presumably due to AO evaporation, appearance of new IR bands were also rapidly detected. In the carbonyls region, the new IR bands were

mainly assigned to quinone methide and cinnamate species, resulting from the chemical consumption of AO. In contrast, in order to explain the appearance of new IR bands in the hydroxyls region, although the polymorphism of Irganox 1076® could not be confirmed by FTIR spectroscopy, a modification of intermolecular interactions involving the phenol functions was also considered.

The FTIR analysis of the thermal ageing in air at 130°C of the homogeneously stabilised Si-XLPE showed the formation of the same oxidation products as observed for the pure AO crystals: quinone methides and cinnamates, then confirmed by UV spectroscopy. For this sample, three main kinetic regimes were observed for AO depletion. In the first regime, AO depletion seems to be mainly due to its physical loss through evaporation, even if some traces of degradation products (quinone methides and cinnamates) resulting from its chemical consumption were also detected. In the second regime, as the AO is still depleting, a broadening of the carbonyl IR massif was observed, mainly due to the formation of oxidation products of AO. Then, in a last regime, IR bands corresponding to oxidation products of the polymer matrix were clearly detected, without any other apparent sign of polymer oxidation. As soon as the phenolic AO was completely depleted (OIT = 0), high concentrations in hydroperoxides were suddenly detected by DSC, whose the immediate decomposition led to a significant auto-acceleration of the thermal oxidation of the polymer matrix.

In the case of the heterogeneously stabilised sample, a continuous decrease in the concentration of AO crystals was observed in the first stage of thermal exposure, until no more AO crystal was detected. This can indicate a further solubilisation of AO into the polymer matrix, as suggested by the surface cleaning treatment (with acetone) performed on some aged samples. This experiment also showed that the oxidation products of AO were both present at the sample surface (i.e. both dissolved and crystallised) and in the sample bulk (i.e. only dissolved). At this stage of investigation, it is, however, still difficult to attribute the chemical consumption observed to a precise phase of AO (i.e. dissolved or crystallised).

In terms of global AO depletion, similar behaviours were observed for both the homogeneously and heterogeneously stabilised samples, which might indicate that AO crystals can act in a similar manner than dissolved AO during thermal ageing. Indeed, while an oxidation induction period at 130°C in air of about 100 days was observed for the homogeneously stabilised sample, it was pushed back to about 400 days for the heterogeneously stabilised sample. This results clearly indicates that AO in excess efficiently participate to the stabilisation of the polymer matrix, as already observed in our previous work [33]. Further investigations would, however, be needed to check if similar results can also be obtained at lower temperatures of exposure, but also to finally propose a general kinetic model for Si-XLPE stabilisation by an excess of Irganox 1076®.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

CRediT authorship contribution statement

Anne Xu: Methodology, Investigation, Conceptualization, Writing - original draft. **Sébastien Roland:** Validation, Writing - review & editing. **Xavier Colin:** Validation, Writing - review & editing, Supervision.

Acknowledgments

This research is part of the European Project “TeaM Cables” aimed at providing to nuclear power plants (NPP) operators a novel methodology for a more efficient and reliable management of the electrical cables ageing, in particular by developing specific experimental and kinetic modelling tools.

The project leading to this application has received funding from the Euratom research and training program 2014–2018 under grant agreement No 755183.

References

- [1] J.W. Martin, R.A. Ryntz, J. Chin, R. Dickie, *Service Life Prediction of Polymeric Materials: Global Perspectives*, Springer Science & Business Media, 2008.
- [2] N. Khelidj, X. Colin, L. Audouin, J. Verdu, A simplified approach for the lifetime prediction of PE in nuclear environments, *Nuclear Instruments and Methods in Physics Research Section B: Beam Interactions with Materials and Atoms* 236 (2005) 88–94.
- [3] N. Khelidj, X. Colin, L. Audouin, J. Verdu, C. Monchy-Leroy, V. Prunier, Oxidation of polyethylene under irradiation at low temperature and low dose rate. Part I. The case of “pure” radiochemical initiation, *Polymer Degradation and Stability* 91 (2006) 1593–1597.
- [4] X. Colin, C. Monchy-Leroy, L. Audouin, J. Verdu, Lifetime prediction of polyethylene in nuclear plants, *Nuclear Instruments and Methods in Physics Research Section B: Beam Interactions with Materials and Atoms* 265 (2007) 251–255.
- [5] X. Colin, E. Richaud, J. Verdu, C. Monchy-Leroy, Kinetic modelling of radiochemical ageing of ethylene-propylene copolymers, *Radiation Physics and Chemistry* 79 (2010) 365–370.
- [6] X. Colin, B. Fayolle, L. Audouin, J. Verdu, About a quasi-universal character of unstabilised polyethylene thermal oxidation kinetics, *Polymer Degradation and Stability* 80 (2003) 67–74.
- [7] X. Colin, L. Audouin, J. Verdu, Determination of thermal oxidation rate constants by an inverse method. Application to polyethylene, *Polymer Degradation and Stability* 86 (2004) 309–321.
- [8] N. Khelidj, X. Colin, L. Audouin, J. Verdu, C. Monchy-Leroy, V. Prunier, Oxidation of polyethylene under irradiation at low temperature and low dose rate. Part II. Low temperature thermal oxidation, *Polymer Degradation and Stability* 91 (2006) 1598–1605.
- [9] X. Colin, B. Fayolle, L. Audouin, J. Verdu, The classical kinetic model for radical chain oxidation of hydrocarbon substrates initiated by bimolecular hydroperoxide decomposition, *International Journal of Chemical Kinetics* 38 (2006) 666–676.
- [10] E. Richaud, X. Colin, B. Fayolle, L. Audouin, J. Verdu, Induction period in the low-temperature thermal oxidation of saturated hydrocarbons: Example of polyethylene, *International Journal of Chemical Kinetics* 40 (2008) 769–777.
- [11] D.W. van Krevelen, K. te Nijenhuis, *Properties of Polymers: Their Correlation with Chemical Structure; their Numerical Estimation and Prediction from Additive Group Contributions*, Elsevier, 2009.
- [12] E. Linde, L. Verardi, D. Fabiani, U.W. Gedde, Dielectric spectroscopy as a condition monitoring technique for cable insulation based on crosslinked polyethylene, *Polymer Testing* 44 (2015) 135–142.
- [13] B. Fayolle, X. Colin, L. Audouin, J. Verdu, Mechanism of degradation induced embrittlement in polyethylene, *Polymer Degradation and Stability* 92 (2007) 231–238.
- [14] J. Verdu, *Vieillessement chimique*, Techniques de l'ingénieur, 1990.
- [15] T. Seguchi, K. Tamura, A. Shimada, M. Sugimoto, H. Kudoh, Mechanism of antioxidant interaction on polymer oxidation by thermal and radiation ageing, *Radiation Physics and Chemistry* 81 (2012) 1747–1751.
- [16] S. Al-Malaika, S. Riasat, C. Lewucha, Reactive antioxidants for peroxide crosslinked polyethylene, *Polymer Degradation and Stability* 145 (2017) 11–24.
- [17] S. Al-Malaika, *Perspectives in Stabilisation of Polyolefins*, in: A.-C. Albertsson (Ed.), *Long Term Properties of Polyolefins*, Springer Berlin Heidelberg, Berlin, Heidelberg, 2004: pp. 121–150.
- [18] J. Pospíšil, Mechanistic action of phenolic antioxidants in polymers—A review, *Polymer Degradation and Stability* 20 (1988) 181–202.
- [19] J. Pospíšil, Transformations of phenolic antioxidants during the inhibited oxidation of polymers, *Pure and Applied Chemistry* 36 (1973) 207–232.
- [20] J. Pospíšil, Chemical and photochemical behaviour of phenolic antioxidants in polymer stabilization—a state of the art report, Part I, *Polymer Degradation and Stability* 40 (1993) 217–232.
- [21] J. Pospíšil, Chemical and photochemical behaviour of phenolic antioxidants in polymer stabilization: A state of the art report, part II, *Polymer Degradation and Stability* 39 (1993) 103–115.
- [22] Y.G. Hsuan, R.M. Koerner, Antioxidant Depletion Lifetime in High Density Polyethylene Geomembranes, *Journal of Geotechnical and Geoenvironmental Engineering* 124 (1998) 532–541.
- [23] H. Zweifel, *Stabilization of Polymeric Materials*, Springer Science & Business Media, 2012.
- [24] S. Al-Malaika, Oxidative degradation and stabilisation of polymers, *International Materials Reviews* 48 (2003) 165–185.
- [25] H. Zweifel, R.D. Maier, M. Schiller, *Plastics Additives Handbook*, Hanser Publications, 2009.

- [26] E. Richaud, J. Verdu, *Vieillessement chimique des polymères - Physicochimie de la stabilisation, Techniques de l'ingénieur*, 2012.
- [27] X. Colin, B. Fayolle, L. Audouin, J. Verdu, *Phénomènes de transport des stabilisants dans les polyoléfines - Partie II : Solubilité, Matériaux & Techniques 91 (2003) 9–14.*
- [28] J.H. Hildebrand, *Solubility of non-electrolytes, Journal of the Society of Chemical Industry 55 (1936) 665–665.*
- [29] A. Xu, S. Roland, X. Colin, *Physico-chemical characterization of the blooming of Irganox 1076® antioxidant onto the surface of a silane-crosslinked polyethylene, Polymer Degradation and Stability 171 (2020) 109046.*
- [30] H.E. Bair, *Exudation of an antioxidant additive from thin polyethylene films, Polymer Engineering & Science 13 (1973) 435–439.*
- [31] J. Saunier, V. Mazel, C. Paris, N. Yagoubi, *Polymorphism of Irganox 1076®: Discovery of new forms and direct characterization of the polymorphs on a medical device by Raman microspectroscopy, European Journal of Pharmaceutics and Biopharmaceutics 75 (2010) 443–450.*
- [32] M. Nouman, J. Saunier, E. Jubeli, N. Yagoubi, *Additive blooming in polymer materials: Consequences in the pharmaceutical and medical field, Polymer Degradation and Stability 143 (2017) 239–252.*
- [33] A. Xu, S. Roland, X. Colin, *Physico-chemical analysis of a silane-grafted polyethylene stabilised with an excess of Irganox 1076®. Proposal of a microstructural model, Polymer Degradation and Stability 183 (2021) 109453.*
- [34] P. Pagès, *Characterization of polymer materials using FT-IR and DSC techniques, Universidade da Coruña, 2005.*
- [35] K. Möller, T. Gevert, *A solid-state investigation of the desorption/evaporation of hindered phenols from low density polyethylene using FTIR and UV spectroscopy with integrating sphere: The effect of molecular size on the desorption, Journal of Applied Polymer Science 61 (1996) 1149–1162.*
- [36] P.P. Klemchuk, P.-L. Horng, *Transformation products of hindered phenolic antioxidants and colour development in polyolefins, Polymer Degradation and Stability 34 (1991) 333–346.*
- [37] A. Xu, S. Roland, X. Colin, *Thermal ageing of a silane-crosslinked polyethylene stabilised with a thiodipropionate antioxidant, Polymer Degradation and Stability 181 (2020) 109276.*
- [38] J. Saunier, J.-M. Herry, N. Yagoubi, C. Marlière, *Exploring complex transitions between polymorphs on a small scale by coupling AFM, FTIR and DSC: the case of Irganox 1076® antioxidant, RSC Advances 7 (2017) 3804–3818.*
- [39] S. Bensason, J. Minick, A. Moet, S. Chum, A. Hiltner, E. Baer, *Classification of homogeneous ethylene-octene copolymers based on comonomer content, Journal of Polymer Science Part B: Polymer Physics 34 (1998) 1301–1315.*
- [40] K. Sirisinha, S. Chimdist, *Comparison of techniques for determining crosslinking in silane-water crosslinked materials, Polymer Testing 25 (2006) 518–526.*
- [41] K. Molt, D. Ihlbrock, *Principles and applications of quality control by near infrared spectroscopy using the example of polymer additives, Fresenius J Anal Chem 348 (1994) 523–529.*
- [42] R. Navarro, L. Audouin, J. Verdu, *Reactions of antioxidants with molecular oxygen. Part II. Isooctyl-3-(3,5-di-tert-butyl-4-hydroxyphenyl)propionate in silicone matrix, Polymer Degradation and Stability 96 (2011) 965–973.*
- [43] R. Navarro, L. Audouin, J. Verdu, *Reactions of antioxidants with molecular oxygen. Part III: Influence of phenolic stabiliser structures on their oxidation in an inert matrix, Polymer Degradation and Stability 96 (2011) 1389–1396.*
- [44] S. Reißmann, M. Stifftinger, K. Grabmayer, G. Wallner, D. Nitsche, W. Buchberger, *Monitoring the degradation of stabilization systems in polypropylene during accelerated aging tests by liquid chromatography combined with atmospheric pressure chemical ionization mass spectrometry, Polymer Degradation and Stability 98 (2013) 1655–1661.*
- [45] J. Pospíšil, W.-D. Habicher, J. Pilař, S. Nešpůrek, J. Kuthan, G.-O. Piringer, H. Zweifel, *Discoloration of polymers by phenolic antioxidants, Polymer Degradation and Stability 77 (2002) 531–538.*
- [46] L.V. Samsonova, L. Taimr, J. Pospíšil, *Antioxidants and stabilizers. LXVI. Oxidation and photooxidation of methyl- and octadecyl-3-(3,5-di-tert-butyl-4-hydroxyphenyl)propionates. quenching of singlet oxygen, Die Angewandte Makromolekulare Chemie 65 (1977) 197–210.*
- [47] C. Berthomieu, E. Nabedryk, W. Mäntele, J. Breton, *Characterization by FTIR spectroscopy of the photoreduction of the primary quinone acceptor QA in photosystem II, FEBS Letters 269 (1990) 363–367.*
- [48] B.H. Stuart, *Infrared Spectroscopy: Fundamentals and Applications, John Wiley & Sons, 2004.*
- [49] A.H. Ali, K.S.V. Srinivasan, *Photoresponsive Functionalized Vinyl Cinnamate Polymers: Synthesis and Characterization, Polymer International 43 (1997) 310–316.*
- [50] N.S. Allen, D. Mallon, A. Timms, A.W. Green, F. Catalina, *Synthesis and spectroscopic properties of novel cinnamate derivatives of benzophenone: Photocuring activity versus photodimerization, European Polymer Journal 29 (1993) 533–538.*
- [51] M.St.G. Flett, *Intensities of some group characteristic infra-red bands, Spectrochimica Acta 18 (1962) 1537–1556.*
- [52] M.D. Cruz, L.V. Schoors, K. Benzarti, X. Colin, *Thermo-oxidative degradation of additive free polyethylene. Part I. Analysis of chemical modifications at molecular and macromolecular scales, Journal of Applied Polymer Science 133 (2016) 43287.*
- [53] E. Richaud, L. Audouin, X. Colin, B. Fayolle, J. Verdu, *Rôle des hydroperoxydes dans le vieillissement oxydant des matériaux polymères, Mater. Tech. 96 (2008) 27–42.*
- [54] B. Peters, S.L. Scott, A. Fong, Y. Wang, A.E. Stigman, *Reexamining the evidence for proton transfers in ethylene polymerization, PNAS 112 (2015) E4160–E4161.*
- [55] J.R. Nielsen, A.H. Woollett, *Vibrational Spectra of Polyethylenes and Related Substances, J. Chem. Phys. 26 (1957) 1391–1400.*
- [56] J. Rud Nielsen, R.F. Holland, *Dichroism and interpretation of the infrared bands of oriented crystalline polyethylene, Journal of Molecular Spectroscopy 6 (1961) 394–418.*
- [57] G.L. Oliveira, M.F. Costa, *Optimization of process conditions, characterization and mechanical properties of silane crosslinked high-density polyethylene, Materials Science and Engineering: A 527 (2010) 4593–4599.*
- [58] Y.-T. Shieh, T.-H. Tsai, *Silane grafting reactions of low-density polyethylene, J. Appl. Polym. Sci. 69 (1998) 255–261.*
- [59] G.S. Ahmed, M. Gilbert, S. Mainprize, M. Rogerson, *FTIR analysis of silane grafted high density polyethylene, Plastics, Rubber and Composites 38 (2009) 13–20.*
- [60] E. Richaud, F. Farcas, B. Fayolle, L. Audouin, J. Verdu, *Accelerated ageing of polypropylene stabilized by phenolic antioxidants under high oxygen pressure, Journal of Applied Polymer Science 110 (2008) 3313–3321.*
- [61] J. Lerchová, L. Kotulák, J. Rotschová, J. Pilař, J. Pospíšil, *Antioxidants and stabilizers. LXI. Photochemical behavior of transformation products of phenolic antioxidants: 4-tert-butylperoxy-2,5-cyclohexadienone, Journal of Polymer Science: Polymer Symposia 57 (1976) 229–235.*
- [62] J. Tireau, *Propriétés à long terme des gaines de polyéthylène haute densité utilisées pour les ponts à haubans, Thèse de doctorat, Arts et Métiers ParisTech (2011).*
- [63] E. Richaud, C. Monchy-Leroy, X. Colin, L. Audouin, J. Verdu, *Kinetic modelling of stabilization coupled with stabilizer loss by evaporation. Case of dithioester stabilized polyethylene, Polymer Degradation and Stability 94 (2009) 2004–2014.*
- [64] F.M. Rugg, J.J. Smith, R.C. Bacon, *Infrared spectrophotometric studies on polyethylene. II. Oxidation, Journal of Polymer Science 13 (1954) 535–547.*
- [65] J.H. Adams, *Analysis of the nonvolatile oxidation products of polypropylene I. Thermal oxidation, J. Polym. Sci. A-1 Polym. Chem. 8 (1970) 1077–1090.*
- [66] D.J. Carlsson, D.M. Wiles, *The Photodegradation of Polypropylene Films. II. Photolysis of Ketonic Oxidation Products, Macromolecules 2 (1969) 587–597.*
- [67] J. Lacoste, D.J. Carlsson, S. Falicki, D.M. Wiles, *Polyethylene hydroperoxide decomposition products, Polymer Degradation and Stability 34 (1991) 309–323.*
- [68] Q. Pelzer, *Étude du vieillissement des isolants synthétiques des câbles moyenne tension "HTA"., These de doctorat, Université Grenoble Alpes (ComUE, 2019).*
- [69] U.W. Gedde, M. Ifwarson, *Molecular structure and morphology of crosslinked polyethylene in an aged hot-water pipe, Polym Eng Sci 30 (1990) 202–210.*
- [70] J.Y. Moisan, R. Lever, *Diffusion des additifs du polyéthylène-V: Influence sur le vieillissement du polymère, European Polymer Journal 18 (1982) 407–411.*
- [71] J.Y. Moisan, *Diffusion des additifs du polyéthylène-I: Influence de la nature du diffusant, European Polymer Journal 16 (1980) 979–987.*
- [72] N.C. Billingham, P.D. Calvert, A.S. Manke, *Solubility of phenolic antioxidants in polyolefins, Journal of Applied Polymer Science 26 (1981) 3543–3555.*