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Nano-films of *carbo*-benzene derivatives: scanning probe microscopy analysis and prospects of use in organic solar cells

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Abstract

Three *carbo*-mer derivatives based on a C₁₈Ph₄ core decorated with two identical electro-active groups X, i.e. two aromatic *carbo*-benzenes (**1** and **2**, X = 4-aniliny) and one pro-aromatic *carbo*-quinoid (the *carbo*-TTF **3**, X = 1,3-dithiol-2-ylidene) were studied through Scanning Probe Microscopies (SPMs). Self-Assembled Monolayers (SAMs) were fabricated (thickness ~160 pm), for the two centrosymmetric representatives **1** and **3**, the organization of which on the HOPG substrate was found to be structure-specific. Electrical/electronic properties of the three *carbo*-mers were determined by using Atomic Force Microscopy (AFM) and its electrical modes: Kelvin Probe Force Microscopy (KPFM) and conductive Atomic Force Microscopy (c-AFM). Measurements of the work function (ϕ) through KPFM result in a $\phi = 5.60$ eV value for **1**, 4.97 eV for **2** and 4.82 eV for **3**. Hole mobility (μ) values extracted from local *I-V* plots by using c-AFM are $15 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **1**, $3 \times 10^{-6} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **2** and $87 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **3**. A concept test of the possible application of *carbo*-mers in self-assembled hole transporting monolayer (SA-HTM), with the view to replacing the most common p-type contact used in organic solar cells (OSCs), PEDOT:PSS, is also reported.

Keywords: *Carbo*-mers, scanning probe microscopy analysis, SAMs as hole transport layer, organic solar cells

1. Introduction

The nanoscale structure of molecules and their local electronic properties can be studied by scanning probe microscopies or SPMs, general abbreviation for various techniques using probe instruments, mainly scanning tunneling microscopy (STM) and atomic force microscopy (AFM), which are ideal to study 2D materials [1, 2]. Based on the interaction between a fine probe tip in proximity to a sample, surface morphological issues, electrical and electronic properties can thus be finely investigated [1, 2]. For instance, STM has been used for imaging the process of self-assembled molecular systems, revealing the parameters involved in monolayer formation: physical parameters (i.e., concentration of solutes, wettability, contact angles, surface tension, temperature) and chemical parameters (i.e., molecular structure, orbital configuration) [3]. On the other hand, AFM is regularly used to characterize morphology of the bulk heterojunction (BHJ) blend in organic solar cells (OSCs) [4]. Besides AFM surface topography measurements, various electrical modes of AFM, such as Kelvin probe force microscopy (KPFM) and conductive-AFM (c-AFM), play an increasing role in understanding the influence of the nanoscale morphology of optoelectronics devices [4]. J. Lee *et al.* [5], used KPFM to study the internal potential distribution of the PCDTBT:PCBM-based OSCs active film, demonstrating the usefulness of this non-invasive approach for quantifying and mapping local differences in electrical performance due to structural heterogeneities. Likewise, by using c-AFM measurements, topography and current maps of perovskite solar cells (PSCs) were compared to analyze how the perovskite grain size affects the device conversion efficiency [6].

Very recently, OSCs have reached a power conversion efficiency (PCE) over 18% [7, 8]. In conventional OSC architecture, the transparent anode electrode is generally coated with a p-type material as hole transport layer (HTL) –the most frequently used HTL being PEDOT:PSS–, which prevents a direct contact between the anode and the active layer, thus reducing electrical charge recombination and maximizing the extraction of holes towards the anode for a higher overall performance [9, 10]. However, PEDOT:PSS has several limitations that can severely compromise

the stability of the devices: for example, the chemical interaction between PEDOT:PSS and ITO (indium-tin oxide, one of the most used transparent electrodes) causes the corrosion of ITO due to acidity, and further, the hygroscopic nature of PEDOT:PSS can reduce the device performance [11]. Therefore, the search for new HTL materials to replace PEDOT:PSS is of growing interest. Recently, graphene, as a 2D material, has gained noteworthy attention because it possesses excellent intrinsic properties, such as high thermal conductivity, transparency, and high electrical-charge mobility [11, 12]. Several groups have used graphene derivatives as HTL in OSCs [11-16]. In this context, self-assembled monolayers (SAMs) of organic molecules, as 2D sheets, have also been considered as an alternative to conventional HTLs [17]. The main advantage of SAMs is the ability to tune the electrical properties of the surface on which they are attached, the modulation of the work function allows a proper energy alignment between the electrode and the active layer [18]. Alternative HTLs based on small molecular materials as SAMs have been used in solar cells showing similar or even higher PCE values than when using PEDOT:PSS [8, 17, 18].

Besides graphene and graphene oxide derivatives, attractive for their transparency and charge carrier properties, other carbon- and electron-rich molecules, namely *carbo*-mers [19], are envisaged for the formation of SAMs. *Carbo*-mers constitute a family of carbon-enriched molecules that are formally obtained by systematic insertion of C₂ units into each (or a selected part) of the covalent bonds for a representative Lewis structure of any parent molecule [19]. Ring *carbo*-mers of benzenic derivatives have been thoroughly studied during the last 20 years, at both the theoretical and experimental levels, and were shown to exhibit unique features making them attractive for several applications [20-22]. In particular, the planar hexagonal π -conjugated C₁₈ macrocycle of *carbo*-benzenes confers them, depending on their substituents, linear and nonlinear optical [23-25], mesogenic or conductive properties [26, 27]. The charge transport properties of *carbo*-benzenes, associated with their flat flexible "*carbo*-aromatic" structure [21], still deserve further investigations, in particular regarding the formation of SAMs to conform novel 2D materials, and their testing as HTL.

In the present work, SPM techniques are used to characterize three *carbo*-meric molecules of different symmetries and self-assembled properties. The first of the tested molecules is a dianilin-4-yl-*carbo*-benzene reported as a centrosymmetric (quadrupolar) representative **1** [28]; the second

is a new dianilin-4-yl-*carbo*-benzene with dipolar structure whose synthesis is herein documented corresponding to the representative **2**; the third molecule is a centrosymmetric *carbo*-quinoid, i.e., 3-dithiol-2-ylidene-*carbo*-quinoid, recently reported and so-called "*carbo*-TTF" **3** [29] (see Fig. 1a). STM images of **1** and **3** show the 2D arrangement of these C₁₈ ring *carbo*-mers on a Highly Oriented Pyrolytic Graphite (HOPG) substrate. Electrical/electronic properties of the three *carbo*-mers were determined by using AFM and its electrical modes: KPFM and c-AFM. Work function (ϕ) values measured through KPFM are $\phi = 5.60$ eV for **1**, 4.97 eV for **2**, and 4.82 eV for **3**. Hole mobility (μ) values extracted from local *I-V* plots by using c-AFM are $15 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **1**, $3 \times 10^{-6} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **2**, and $87 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **3**. In addition, the three *carbo*-mers were implemented (concept tests) for the formation of possible self-assembled hole-transport monolayers (SA-HTL) in OSCs. OSCs were fabricated under the direct configuration: ITO/HTL/PEDOT:PSS/PBDB-T:ITIC/PFN/FM, where PBDB-T is the electro-donor, ITIC is a non-fullerene acceptor, PFN is an alcohol/water soluble conjugated polymer used as electron transport layer (ETL) [30], see Fig. 1a); and FM = Field's metal is an alternative top electrode composed of Bi, Sn and In that has been previously reported by our group [30-32]. The best PCE with a HTL of PEDOT:PSS was 5.5%, whereas with **2** as SA-HTM it was 3.9%. Without a HTL, OSCs showed a PCE of 2.7%. Although PCE values close to 9 % have been achieved in our group under the same configuration (with PEDOT:PSS as HTL) [2, 16], for this current study the maximum PCE was 5.5% because of the non-optimized procedure followed in [2, 16]; however, the main purpose in these concept tests is just to compare PV performance of OSCs with PEDOT:PSS and *carbo*-mer SAMs as HTL. These results show that upon optimization, *carbo*-mer-based SA-HTLs would be possible alternatives for OSCs.

2. Experimental

2.1 Material and reagents

Poly[(2,6-(4,8-bis(5-(2-ethylhexyl)thiophen-2-yl)-benzo[1,2-b:4,5-b']dithiophene))-alt-(5,5-(1',3'-di-2-thienyl-5',7'-bis(2-ethylhexyl)benzo[1',2'-c:4',5'-c']dithiophene-4,8-dione)] (PBDB-T), 3,9-bis(2-methylene-(3-(1,1-dicyanomethylene)-indanone))-5,5,11,11-tetrakis(4-hexylphenyl)-dithieno[2,3-d:2',3'-d']-s-indaceno[1,2-b:5,6-b']dithiophene (ITIC) and, the alcohol/water-soluble conjugated polymer poly[(9,9-bis(3'-(N,N-dimethylamino)propyl)-2,7-fluorene)-alt-2,7-(9,9-dioctylfluorene)] (PFN), were acquired from 1-Material Inc. The eutectic

alloy Field's metal (FM, 32.5% Bi, 51% In, 16.5% Sn) was obtained from Rotometals and ITO-covered glass substrates (10 Ω /square) from Delta Technologies. 1-Phenyloctane was purchased from Sigma-Aldrich and used without further purification.

Tetrahydrofuran (THF), diethyl ether (Et₂O), pentane and dichloromethane (DCM) used for the synthesis of the *carbo*-mers **1**, **2**, **3** were dried with a PureSolve-MD-5 Innovative Technology system for the purification of solvents. All other reagents were used as commercially available. In particular, the commercial solution of HCl was 2 M in diethyl ether, and SnCl₂ was labelled anhydrous. Silica gel (60 Å, C.C. 70-200 μ m) was used for column chromatography. Silica gel thin layer chromatography (TLC) plates (60F254, 0.25 mm) were revealed under UV-light and/or by treatment with an ethanolic solution of phosphomolybdic acid (20 %). The following ¹H and ¹³C NMR spectrometers were used: DPX 300, Avance 500. The ¹H and ¹³C NMR signals were assigned according to chemical shifts, spin-spin coupling constants, splitting patterns, signal intensities; NMR chemical shifts δ are given in ppm with positive values at high frequency relative to the tetramethylsilane reference. Coupling constants *J* are in hertz.

2.2 Synthesis of the *carbo*-benzene **2**

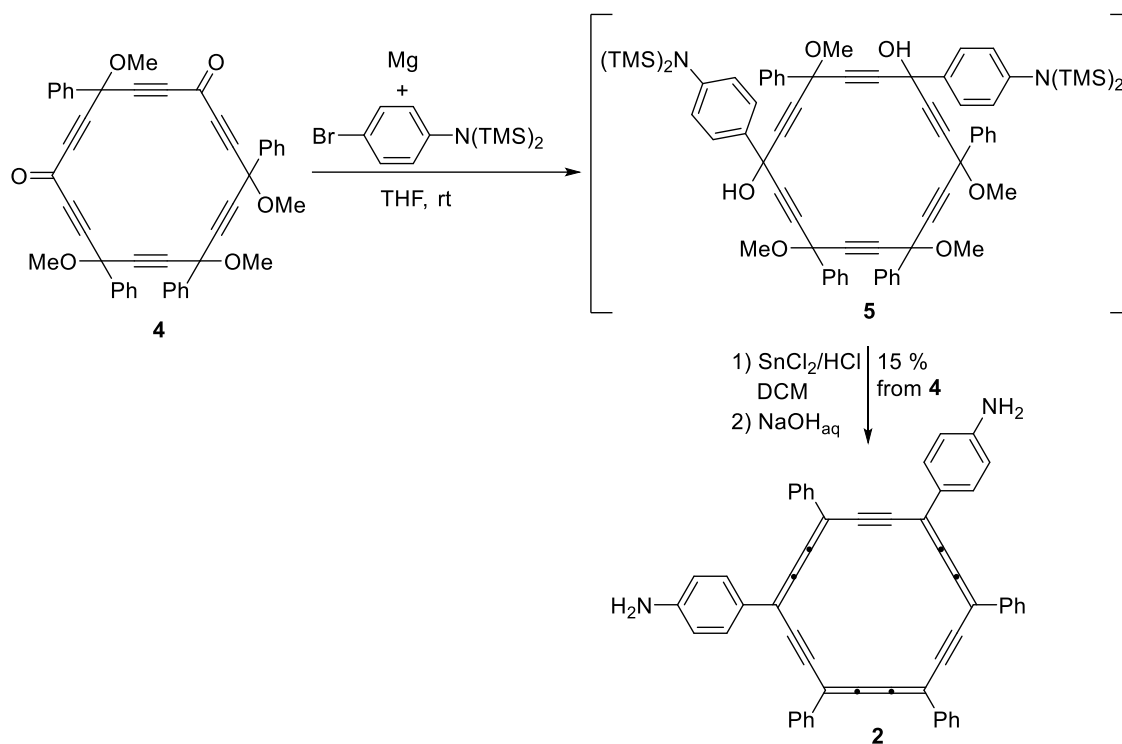
2 was prepared in two steps from the known [6]pericyclinedione **4** (Scheme 1) [33].

To a solution of **4** (0.100 g, 0.147 mmol) in THF (10 mL) was added at 0 °C a solution of Grignard reactant, itself prepared by addition of bromo-bis-trimethylsilylaniline (0.23 mL, 0.88 mmol) to Mg (0.021 g, 0.86 mmol) in THF (1.5 mL) at room temperature. The resulting mixture was stirred 1 h at 0 °C and overnight at room temperature, before treatment with water. The aqueous layer was extracted with diethyl ether, and the combined organic layers were washed with brine, separated, dried with MgSO₄ and concentrated under reduced pressure. The residue containing the [6]pericyclinediol **5** was directly engaged in the next step without further purification.

To solution of **5** (0.147 mmol) in DCM (70 mL) under stirring at -78 °C were added successively SnCl₂ (0.283 g, 0.149 mmol) and HCl (1.54 mL, 0.3 mmol). The resulting mixture was allowed to slowly warm up to 0 °C before stirring for 1 h at room temperature. Then an aqueous solution of NaOH (1M, 3 mL) was added and the stirring was maintained overnight. After washing with brine, the organic layer was separated, dried with MgSO₄ and concentrated under reduced pressure. The

residue was purified by silica gel chromatography (pentane/acetone/THF 8:1:1, then 6:2:2 then THF/acetone 1:1) to give **2** as a violet solid (0.016 g, 15 % from **4**).

^1H NMR (THF- D_8) δ 5.67 (broad s, 4H), 7.30 (d, $J = 7.4$ Hz, 4H), 7.80 (m, 4H), 8.10 (m, 8H), 9.32 (d, $J = 7.4$ Hz, 4H), 9.53-9.60 (m, 8H). $^{13}\text{C}\{^1\text{H}\}$ NMR (THF- D_8) δ 99.6, 11.4, 102.6, 104.0, 107.2, 115.1, 115.3, 115.7, 116.5, 117.0, 117.3, 118.6, 128.4, 128.6, 128.8, 129.2, 129.4, 129.5, 129.6, 129.7, 129.9, 132.0, 140.2, 140.4, 140.7. HRMS (MALDI-TOF) m/z calculated for $\text{C}_{54}\text{H}_{33}\text{N}_2$ $[\text{M}+\text{H}]^+$: 709.2644, found : 709.2641.



Scheme 1. Synthesis of the dipolar *carbo*-benzene **2** from the known [6]pericyclynedione **4**.

2.3 Thin films fabrication for SPM measurements

For STM measurements at the liquid/solid interface [34, 35], *carbo*-mer samples (**1**, **2** or **3**) were dissolved in 1-phenyloctane with a concentration about 0.1 mM and then deposited on HOPG by the drop casting method (thickness ~ 160 pm, SAMs). For KPFM and c-AFM determinations, they were dissolved in chlorobenzene with a concentration about 1 mg/mL and then deposited on ITO surface by drop coating technique (thickness ~ 10 nm, thin films).

2.4 Films characterization

STM/AFM topographies were obtained with a Nanosurf Easyscan 2 STM/AFM operating at constant current mode. AFM electrical measurements were carried out by using a SmartSPM-1000 AFM from AIST-NT (current noise of 200 fA). Platinum-coated AFM tips (HQ:NSC15/PT from MikroMasch, PtSi-CONT/NCH and ATEC-CONTPt-SPL from Nanosensors) were used. For KPFM measurement PtSi AFM tips and two-pass mode were used. To calibrate the KPFM system, contact potential difference (CPD) was used where ϕ of PtSi tip was measured by calibrating it with Highly Oriented Pyrolytic Graphite (HOPG) as reference sample, then ITO reference sample was measured to further check this calibration. Before calibration, HOPG substrate was cleaved by using the adhesive tape technique to achieve an atomically clean surface and, ITO substrate was cleaned sequentially in a soap solution, distilled water, acetone, and isopropanol, in each case for 15 min in an ultrasonic bath and dried in an oven at 80 °C. The work function for PtSi was $\phi = 5 \pm 0.13$ eV, which agrees with previous reports [37, 38]. Once the PtSi tip and KPFM system were calibrated, measurements were consistent on the multiple samples with a standard deviation of ± 0.1 eV under the same measurement conditions. The lift height from the two-pass mode was kept around 15-25 nm and an applied potential of ± 4 V. PtSi tip parameters are: force constants $k \approx 0.2$ N/m and a resonant frequency of $f \approx 13$ kHz. Preparation of STM tips and substrates as well as additional technical information can be found elsewhere [2, 35]. The tunneling conditions of each STM image are given in the corresponding Figure caption. Images were processed and analyzed with the Gwyddion open-source software (v. 2.49), from which mean plane subtraction, Gaussian filter, autocorrelation function (ACF), fast Fourier transform (FFT) and measured distance were determined.

2.5 OSCs fabrication and characterization

The ITO-covered glass substrates were cut (~ 1.8 cm \times 1.8 cm) and ultrasonically cleaned consecutively for 20 min in a detergent solution, distilled water, and isopropanol, and dried in an oven at 80 °C for at least 12 h. ITO substrates were treated with oxygen plasma for 20 min before coating. For the HTL, PEDOT:PSS and possible *carbo*-mer SA-HTLs (from **1**, **2** and **3**) were tested (see Fig. 1b). A PEDOT:PSS layer of ~ 40 nm thickness was spin-coated at 4500 rpm for 1 min on top of the ITO substrate and dried in an oven for 30 min at 80 °C. *Carbo*-mers were

dissolved in 1-phenyloctane (c.a. 0.1 mM) and kept under stirring for 30 min. ITO substrates were immersed in SAM solutions for 12 h. After that, substrates were rinsed with toluene to remove residual SAM molecules that would not be strongly bonded on the ITO surface. For the active layer, PBDB-T was used as electro-donor and ITIC as electro-acceptor (see Fig. 1a). The solution of the active layer was prepared by dissolving PBDB-T and ITIC at 1:1 w/w, 20 mg/mL in anhydrous chlorobenzene/1,8-diodooctane (99.5:0.5 v/v) within a glovebox under nitrogen atmosphere. The solution was stirred in a magnetic stir plate for about 24 h at room temperature. Then, it was spin-coated, at 2500 rpm for 1 min (active layer thickness ~ 100 nm) inside the glovebox, onto the PEDOT:PSS or **1**, **2** or **3** monolayer. The active layer was annealed at 160 °C for 10 min. A PFN layer (~5-10 nm) was spin-coated at 6000 rpm on top of the active layer and exposed to thermal annealing for 15 minutes at 80 °C. The active area of devices (0.07 cm²) was delimited with a Scotch tape. FM top electrode was deposited by drop casting on top of the PFN layer, after melting it at 95 °C in a hot plate; during this procedure, the OSC device was kept on the hot plate at the same temperature. Additional details can be found elsewhere [30-32]. Fig. 1a) shows the chemical structure of the studied *carbo*-meric molecules. In Fig. 1b), the glass/ITO/HTL/PBDB-T:ITIC/PFN/FM OSC architecture used in this work is schematized, where HTL is PEDOT:PSS or *carbo*-meric SAMs as mentioned above. The OSCs were illuminated with the Air Mass 1.5 spectrum (1000 W/m²) provided by a solar simulator Sciencetech SS150 class AAA, which was calibrated using a 15151 calibrated reference cell from abet technologies. Current density vs voltage (*J-V*) curves were acquired by a Keithley 2450 source-meter in ambient conditions.

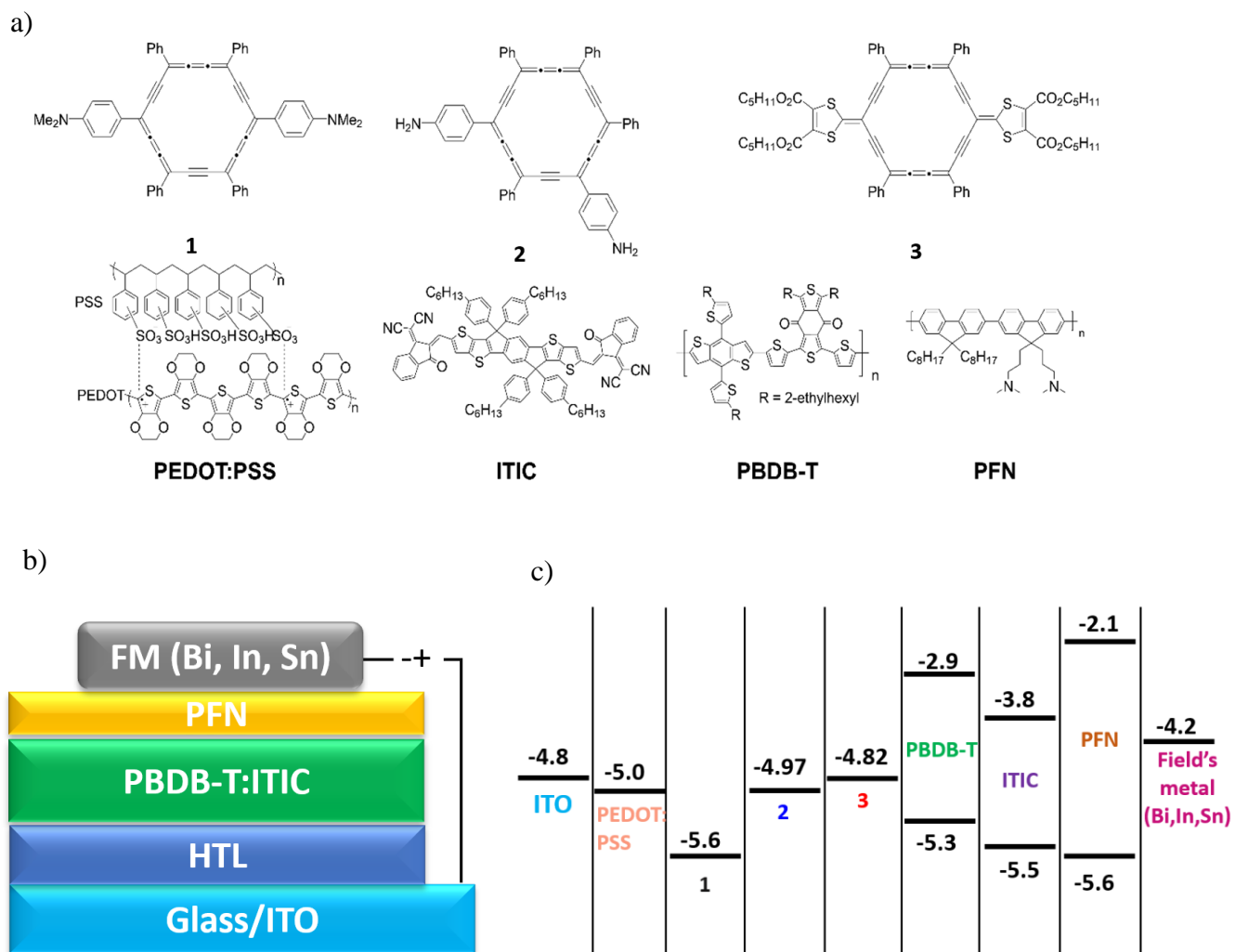


Fig. 1. a) Chemical structures of the compounds used in this work, b) Architecture of the OSC devices. c) Energetic levels of the compounds used in this work (data in eV). Note that our scheme does not intend to represent the absolute values [9-11].

3. Results and Discussion

The optical absorption in solution of the *carbo*-mers **1-3** is displayed in Fig. 2. These spectra present the characteristic features of *carbo*-benzenes [23]; for instance, the *carbo*-benzene **1** bearing NMe₂ substituents on the phenyl groups in *para* position of the C₁₈ macrocycle exhibits an intense absorption band at 2.37 eV (523 nm) and a low-lying energy band of weak intensity at 1.91 eV (649.1 nm). The *carbo*-benzene **2** bearing NH₂ substituents on the phenyl groups in *meta*

positions of the C₁₈ ring has its maximum of absorption at 2.51 eV (494 nm), shifted by +0.14 eV with respect to **1**, characterized as a single main band with a long tail of absorption structured with two shoulders (2.04 and 2.28 eV). Finally, the *carbo*-quinoid **3** has two intense bands in the red and blue regions of the spectrum at 1.96 eV (632.5 nm) and 2.78 eV (446 nm), respectively, resembling the characteristics of the dichromic *carbo*-cyclohexadiene [23]. The inset in Fig. 2 shows the Tauc plots from which optical bandgap energies E_g of 1.82, 2.13 and 1.86 eV are calculated for **1**, **2** and **3**, respectively. Here the optical bandgap is interpreted as the onset of optical absorption that produces excitons. The value of E_g in **2** is due to a combined effect of the substituents on the phenyl groups in *meta* or *para* positions of the *carbo*-benzene and the strength of the donating groups; the E_g value of the *carbo*-quinoid **3** is similar to the E_g value of **1**. Notice that the Tauc plot shows additional linear edges for **1**, **2** and **3**, corresponding to bandgaps of 2.25, 2.33 and 2.64 eV, respectively. Two linear edges in Tauc plots are also commonly observed in other organic materials characterized by structured absorption spectra, such as porphyrins [39-41] where the optical gap of lowest energy is the onset of absorption while the one of higher energy is the fundamental energy gap, namely, the energy gap between a combination of the HOMO π -orbitals and LUMO π^* -orbitals. The energies of 2.25, 2.33 and 2.64 eV are assigned to the fundamental bandgaps of **1**, **2** and **3**. It is thus not surprising then that *carbo*-benzenes can exhibit various bandgaps as their absorption features (various bands denoting electronic transitions) resemble to some extent the Q and B (Soret) bands displayed by porphyrins. The absorption features of *carbo*-benzenes were indeed previously explained in terms of the Gouterman four-orbital model originally developed for porphyrins [42].

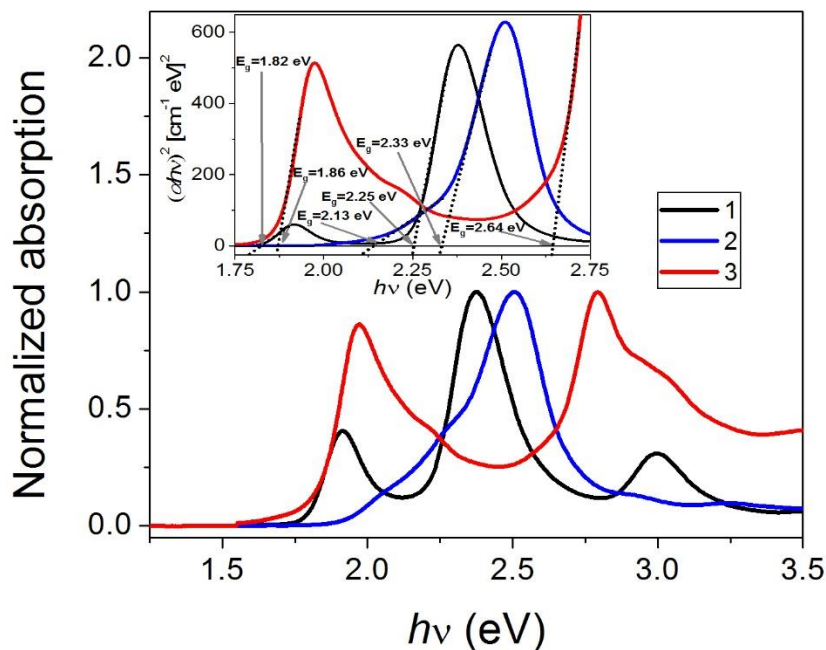


Fig. 2. Absorption spectra of **1**, **2** and **3** in dichloromethane solution. Inset: Tauc plots with dotted lines depicting the linear portion of the spectra from which the optical band-gap E_g is calculated.

Highly ordered self-assembled monolayers of the *carbo*-benzene **1** and *carbo*-quinoid **3** are shown in Figs. 3a)-b). For the *carbo*-benzene **2**, no STM image was observed in 1-phenyloctane at any concentration [43], which does not necessarily mean that a SAM layer is not formed. It has been reported [44, 45] that hydrogen bond donors such as carboxylic acid (CO₂H) or thiol (SH) groups interact with ITO substrates. The *carbo*-benzene **2** can thus also be anticipated to bind to the ITO by its anilinyll amine groups (NH₂) [45]. In the case of the *carbo*-benzene **1** and *carbo*-TTF **3**, a charge transfer interaction with the surface oxide (ITO^{δ-... 1/3}δ⁺) can be invoked to be facilitated by weak oxidation potentials through (quasi)reversible processes ($E_{1/2}(\mathbf{1}) = + 0.51 \text{ V vs SCE}$, $E_{p/1/2}(\mathbf{3}) = + 0.18 \text{ V vs Fc/Fc}^+$, in dichloromethane solution) [28, 29]. Besides, the formation of regular monolayers is expected from the tendency of the macrocyclic cores to stack face-to-face by π - π interactions when titled at surfaces [26, 29]. STM images of **1** and previously described images of **3** [29] (Figs. 3a)-b)) show ordered monolayers along two horizontal (in-plane)

orientations. The polarity of the electron-donating dimethylaniliny groups of **1** plays a key role in the molecular orientation on the substrate, because it leads to a complex combination of forces (attractive and repulsive, short- and long-range) that allows the control of assembly dimensionality on the surface with a dipole moment along the molecular axis [46]. Indeed, all atoms are arranged in such a way that the unit-cell axes of the packing arrangement remain unchanged within the domain itself. Benneckendorf *et al.* [47] reported the structure-property relationships of dipolar phenyl and biphenylphosphonic acid: the strengths of the dipoles are varied through the inclusion of electron-withdrawing fluorine atoms and the molecular length, however, the surface energy is kept constant through the electron-donating *N,N*-dimethylamine fragment, which were able to form self-assembled monolayers on indium-tin oxide (ITO) electrodes modifying its work function (ϕ). In addition, quantum mechanical calculations showed that the dipole vector of *N,N*-dimethylaniline derivatives is orientated towards the electrode interface, directly influencing changes of the ϕ value [47]. Meanwhile, by removing the methyl groups from the amino fragment, from NMe_2 in **1** (quadrupolar) to NH_2 in **2** (dipolar), the work function of the metal surfaces has been modified by the adsorption of SAMs, which form surface dipoles [48], as observed for **1** but also for the dithiolylidene-substituted *carbo*-quinoid **3**. The ACF (from the Gwyddion software), exposes the lattice vectors of the self-assembly and the separation distance between neighboring molecules (see Figs. 3c-d)). Characteristic lattice vector data, through STM, are summarized in Table 1.

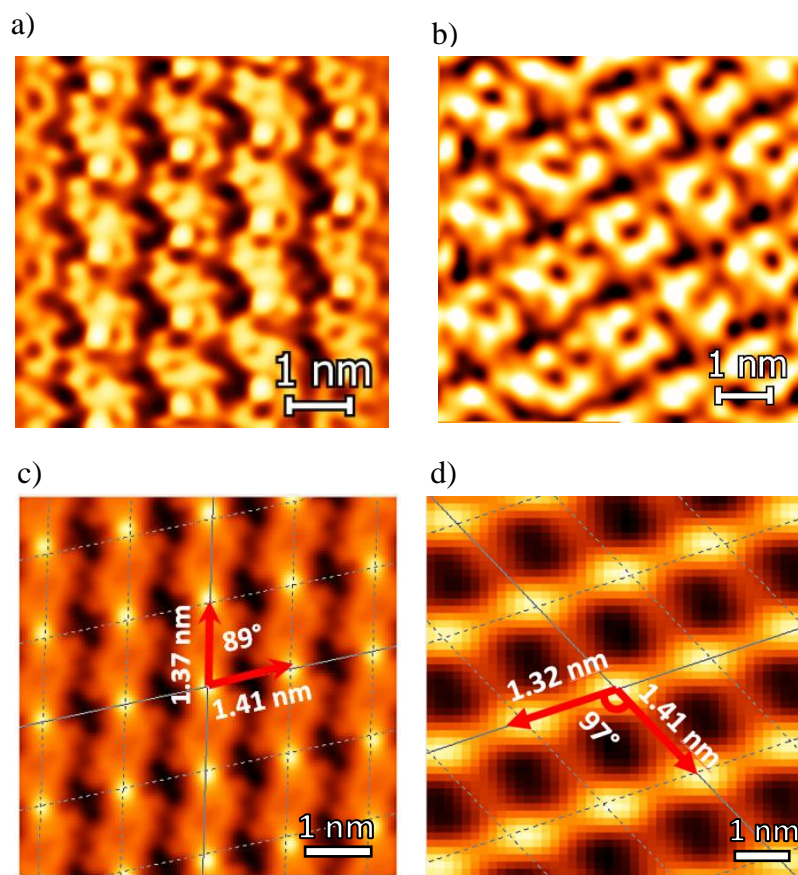


Fig. 3. SAMs of a) **1**, tunneling condition; $I_t = 160$ pA, $U = 0.90$ V, b) **3**, tunneling condition; $I_t = 140$ pA, $U = 1.05$ V. Lattice parameters by using ACF of c) **1**, d) **3**.

Table 1. Lattice parameters of SAMs of the *carbo*-benzene **1** and *carbo*-quinoid **3** on HOPG (STM determinations).

<i>Carbo</i> -mer	a (nm)	b (nm)	α ($^\circ$)
1	1.41	1.37	89
2	–	–	–
3	1.32	1.41	97

As previously reported [29], STM topography of the *carbo*-quinoid **3** exposes extended layers of self-assembled molecules with an average thickness of ~ 160 pm. In the case of the *carbo*-benzene

1, a large scale STM image (Fig. 4a) reveals a columnar arrangement. However, high resolution STM image (Fig. 4b) shows that the macrocycles are twisted out of the molecular plane, which may be attributed to rotation of the already twisted C-N bonds induced by repulsion between the lone-pair electrons at the nitrogen and the substrate (see blue arrows in Fig. 4b) [49], where the bright points can be assigned to the dimethylaniliny groups. Fig. 4c) shows a schematic drawing of the self-assembly of the blue arrows zones in Fig. 4b).

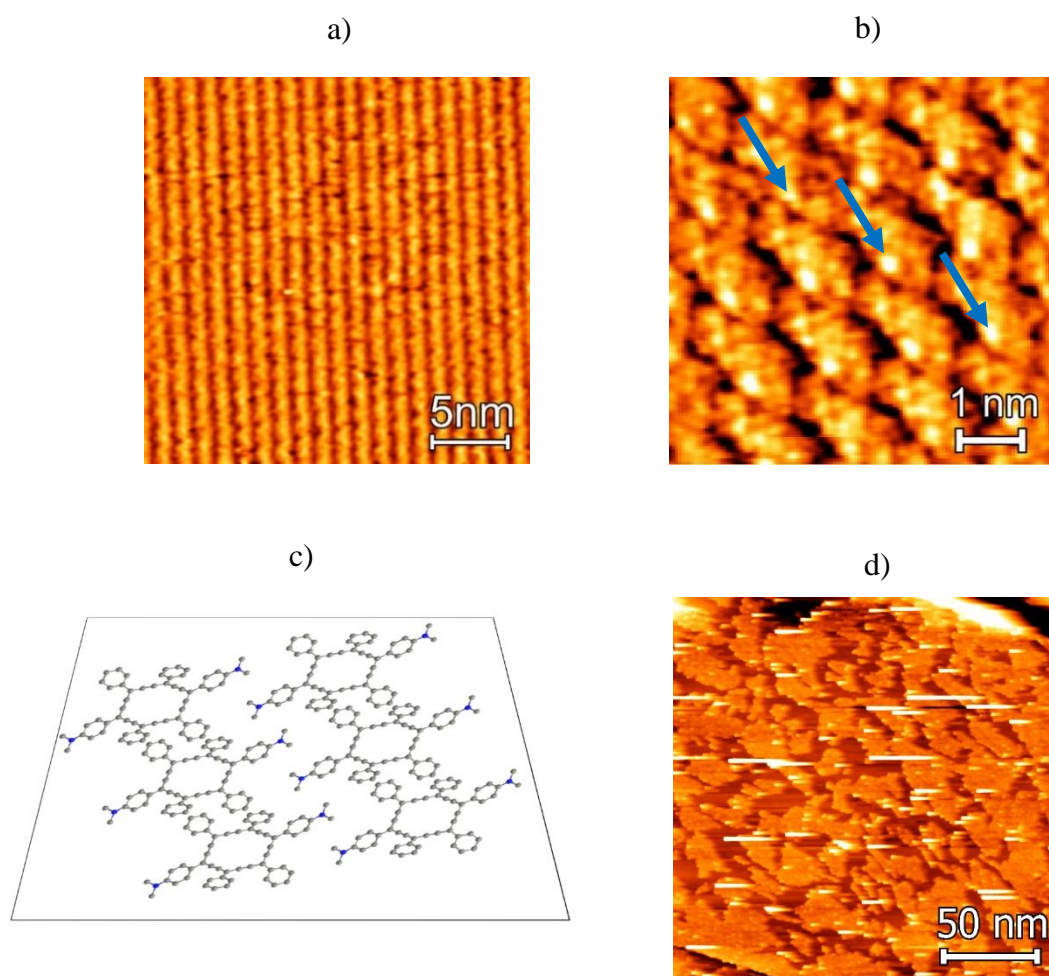


Fig. 4. a) STM image of a self-assembled film of the *carbo*-benzene **1** ($I_t = 140$ pA, $U = 1.05$ V), b) High resolution image ($I_t = 160$ pA, $U = 0.90$ V), c) Schematic arrangement of assembled molecules of **1** (where the blue arrows in b) indicate), d) Large-scale STM images showing the clusters formed by the *carbo*-benzene **1** (horizontal stripes are noise due to unsteady feedback loop –tip oscillation– at large scale scanning).

As suggested by the STM images, the organization of the molecules on the substrate to form SAMs is governed by weak intermolecular interactions between the π -electron-rich, planar and moderately rigid *carbo*-aromatic C₁₈ cores [21], depending on both their substitution pattern (dipolar or quadrupolar) and the nature of their substituents: the molecules are anchored to the substrate but their mean planes are not parallel to its surface, tilting and overlapping with neighbors by π - π stacking. These π -stacking interactions are however likely too weak by themselves to allow the formation of stable multilayers on the ITO without assistance of hydrogen bond donor or acceptor. Indeed, π - π stacking between *carbo*-benzene rings are quite rare, even in the crystal state, with only partial parallel overlap of C₁₈ hexagons, e.g. by 25 % only in titled C₁₈-C₁₈ columnar arrangements of *carbo*-benzene hydrocarbons [26, 50].

For imaging the electrical properties of thin layers of *carbo*-mers, films with a thickness of ~10 nm were deposited on ITO substrates. Figs. 5 a)-c) correspond to a Root Mean Square (RMS) roughness (R_q) analysis resulting in the values: $R_q = 12$ nm for **1**, $R_q = 18$ nm for **2** and $R_q = 8$ nm for **3**; these relatively high roughness values are due to the low solubility of the studied *carbo*-mers, preventing the preparation of highly regular and smooth thin solid films by drop casting or spin casting techniques. Such a lack of solubility of *carbo*-benzene derivatives is a general issue that was shown to prevent their study for organic photovoltaic applications – as either donor or acceptor components in the active layer – as mentioned in previous reports [51, 52].

KPFM is an electrical mode of AFM, which can visualize the changes in the internal potential distribution between the tip and the surface, commonly referred to as the contact potential difference (CPD) [2, 53]. KPFM is one of the most sensitive techniques for ϕ measurement, determined from surface contact potential differences (V_{CPD}) between the conductive tip and the sample [54]. Figs. 5 d)-f) show the potential distribution on the surface of films of *carbo*-mers on ITO measured through KPFM. Here, the use of uniform PtSi-coated tips avoid topography correlated artifacts (topography inhomogeneities induced by electrostatic forces through the contact of the tip with the sample) [55] and the contact potential difference gives ϕ value of 5.60 eV for **1**, 4.97 eV for **2**, 4.82 eV for **3**.

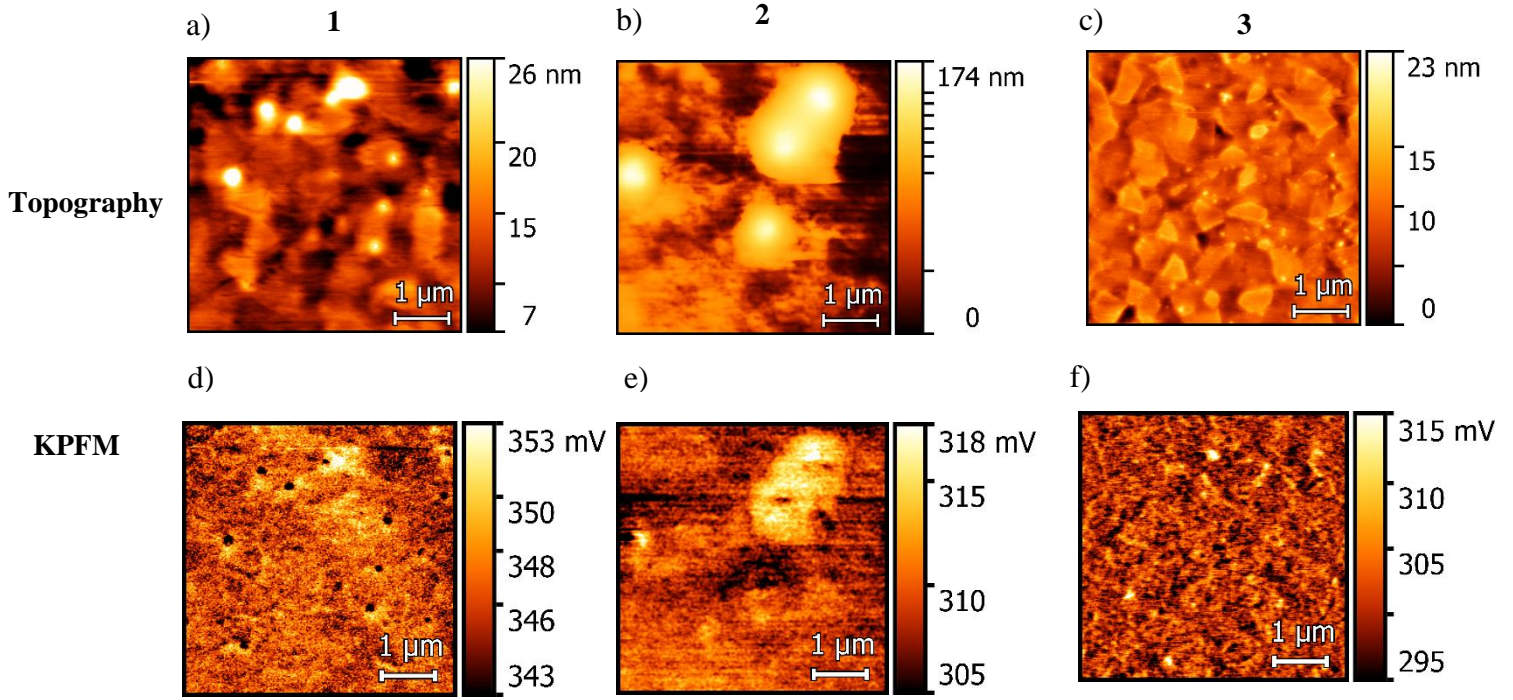


Fig. 5. AFM measurements on thin layers ($T = \sim 10$ nm) of the C_{18} ring *carbo*-mers **1**, **2** and **3**. a), b), c): topography using contact mode AFM; d), e), f): contact potential difference using KPFM.

Fig. 6 displays local I - V curves (linear and root square plots) for the three *carbo*-mer thin films, where each curve being the average of 4 measurements resulting in a relative standard deviation of the mean from 10% – 17 % due to the roughness in each film [56]. From these local I - V curves (Fig. 6b), hole mobility (μ) was extracted by using a modified space-charge limited current (eq. 1) (SCLC) model that accounts for the AFM tip-sample geometry [56, 57]; additional details of the used parameters for the quantitative interpretation can be found elsewhere [2, 56, 57].

$$J = 8.2 \varepsilon_o \varepsilon_r \mu \frac{V^2}{T^3} \delta \left(\frac{T}{d} \right)^{1.6 \pm 0.1} \quad (1)$$

where ε_o and ε_r are the vacuum dielectric permittivity constant and the thin film dielectric constant (relative electrical permittivity), respectively; for organic semiconductors it was assumed $\varepsilon_r = 3$ [57], μ is the charge carrier mobility, T is the thickness of the semiconductor thin film and d is the tip-contact radius. For the contact radius d , the tip-substrate contact area was estimated from DMT model [56] (for PtSi probes the contact radius d is 5.8 nm). The constant 8.2 is the prefactor determined from the finite element modeling to account for the nonuniform electrical field [57].

The constant $\delta = 7.8 \pm 1$ was included to empirically account for the difference between mobilities derived from c-AFM measurements and those from planar electrodes. Average local film thickness is determined by tapping mode AFM, where topographic height (roughness) at each sample location is measured together with the film thickness. However, the accuracy of thickness measurement of thin films is limited when the surface is enough rough [58]. Therefore, the hole mobility is thickness dependence ($J \propto T^{-1.4 \pm 0.1}$) when carrying out the experiment using c-AFM, as well as dependence in the tip diameter (geometry) and current fluctuations due to film roughness. The above-mentioned dependencies lead that to the estimated hole mobility of sample **3** be larger than that of sample **1** even though slope in Fig. 6b is larger for **1** (see Table 2).

The I - V curves (Fig. 6a) show a non-ohmic or “S” shape behavior, suggesting that charge transport occurs through a tunnel potential barrier, although the responses are different in the three thin films [48] (**3** exhibits a wider and flat curve – built-in potential from the y-intercept– than **1** and **2**): higher current is observed when the slope is greater [59]. The estimated hole mobility values are $15 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **1**, $3 \times 10^{-6} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **2** and $87 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **3**. These values are lower than mobility of PEDOT:PSS, which is in the range of $3 \times 10^{-4} - 5 \times 10^{-5} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ [60, 61]. Table 2 summarizes the morphology, electrical and electronic properties of the three *carbo*-mers.

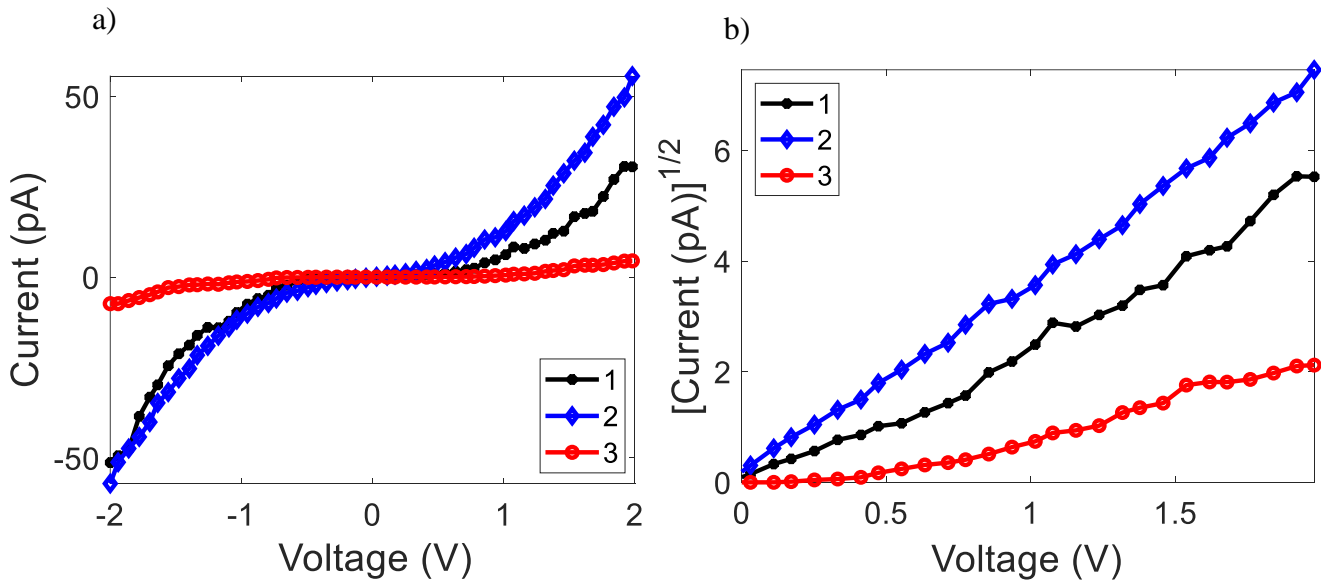


Fig. 6. Local I - V plots acquired by c-AFM, for thin films (~ 10 nm) of the C_{18} ring *carbo*-mers **1**, **2**, **3**, a) Linear scale, b) $I^{1/2}$ - V plot.

Table 2. Morphology, electrical and electronic properties of thin films of the C_{18} ring *carbo*-mers **1**, **2**, **3**, using chlorobenzene as solvent, for thin films of *ca* 10 nm thickness (PEDOT:PSS parameters are also shown for comparison purposes).

<i>Carbo</i> -mer	Roughness (nm)	Work function (ϕ) (KPFM) (eV)	Hole mobility (μ) (c-AFM) ($cm^2V^{-1}s^{-1}$)	Optical Bandgap (eV)	Reference
1	12	5.60	15×10^{-8}	1.82	This work
2	18	4.97	3×10^{-6}	2.13	This work
3	8	4.82	87×10^{-8}	1.86	This work
PEDOT:PSS	2	5.0 ^a	$3 \times 10^{-4} - 5 \times 10^{-5}$ ^a	2–2.2	[60, 61]

^a WF for PEDOT:PSS was determined by KPFM and hole mobility through space-charge-limited current (SCLC) technique with the hole-only devices fabricated.

The use of SAMs of *carbo*-mers as HTLs could offer the benefits of uniform layers of minimal thickness (~ 160 pm). Fig. 7 shows preliminary results with the C_{18} ring *carbo*-mers **1**, **2** or **3** as SA-HTL in OSCs. As can be seen from Fig. 7 and Table 3, with these SA-HTLs, OSCs show lower efficiency than by using PEDOT:PSS as HTL, probably because the low concentration of *carbo*-mer solutions at 0.1 mM in 1-phenyloctane (Fig. 3a) did not allow a full homogeneous coverage of the ITO substrate by the SAMs, resulting in low ohmic contact with the active layer [16]. For this reason, the preparation of the SAMs needs to be optimized to improve the efficiency of related *carbo*-mers in OSCs. Optimization of concentration, solvent (or mixtures of solvents) and drying time could allow obtaining a uniform monolayer homogeneously covering the ITO surface, and thus possibly reaching higher efficiencies. However, these J - V plots, as test of concept, evidence the possibility to use *carbo*-mers as SAMs for replacing PEDOT:PSS as HTL. The best PV performance was determined with a SAM based on **2**, perhaps because of its larger

hole mobility value (see Table 2) than that measured for thin films of **1** and **3**. The low mobility of *carbo*-mers prevents efficient charge extraction in OSCs, which need a hole mobility ca $10^{-3} \text{ cm}^2\text{V}^{-1}\text{s}^{-1}$ [60, 61]. Another factor that can affect this PV behavior is the estimated ϕ value of 4.93 eV for **2**, which is lower than that for **1**: 5.60 eV (however, slightly higher than that for **3**: 4.82 eV), and its matching with that of ITO: 4.8 eV. Since a higher ϕ value than that of ITO will provide a barrier to hole extraction, it is desired that ϕ is close to the active layer's highest occupied molecular orbital (HOMO) level (PBDB-T; 5.3 eV) [62], and then hole injection from the active layer to the anode is facilitated. The above-indicated HOMO levels and ϕ values reflect the relative positions of the transport energy levels (absolute values). Furthermore, according to Fig. 2 and Tauc plot, the largest E_g value of the three *carbo*-mers is precisely that of **2**: 2.13 eV. The wide E_g for **2** and its ϕ value could better allow to block electrons, thus preventing large energy loss resulting from a band offset between the HOMO level of PBDB-T (5.21 eV) and the SAM [63]. Nevertheless, these PCE values are still lower than those reported when PEDOT:PSS is used as HTL mainly because the well optimized parameters of PEDOT:PSS (see Table 2), thus requiring future optimization of both the preparation of the SAMs and possibly to the structure of *carbo*-mers. However, due to the well-known hydrophilic and acidic nature of PEDOT:PSS [13], it is expected that its replacement by more hydrophobic components would be a first step to improve the device stabilization. Several materials have been previously reported as alternative to PEDOT:PSS to improve stability and in some cases to achieve a higher PCE value, such as SAMs [8, 17, 18] and layers of graphene derivatives [11]. While such sp^2 -carbon-rich aromatic materials thus tend to repel water molecules from the device, the effect is expected to be at least similar for the present sp -/ sp^2 -carbon-rich *carbo*-aromatic molecules [21]. It is also noteworthy that a *carbo*-benzene bearing ethynylanilinyll substituents was recently shown to form a stable hydrophobic layer on another metallic oxide (TiO_2) covered with silver nanoparticles (Ag) and to stabilize these two components in water suspensions, that are *a priori* oxidizing (as being used for photocatalytic water splitting: $2 \text{ H}_2\text{O} \rightarrow 2 \text{ H}_2 + \text{O}_2$), over long periods [24, 64]. Furthermore, the ability to form monolayers makes *carbo*-mers promising candidates for reducing manufacturing costs since their extremely low material consumption.

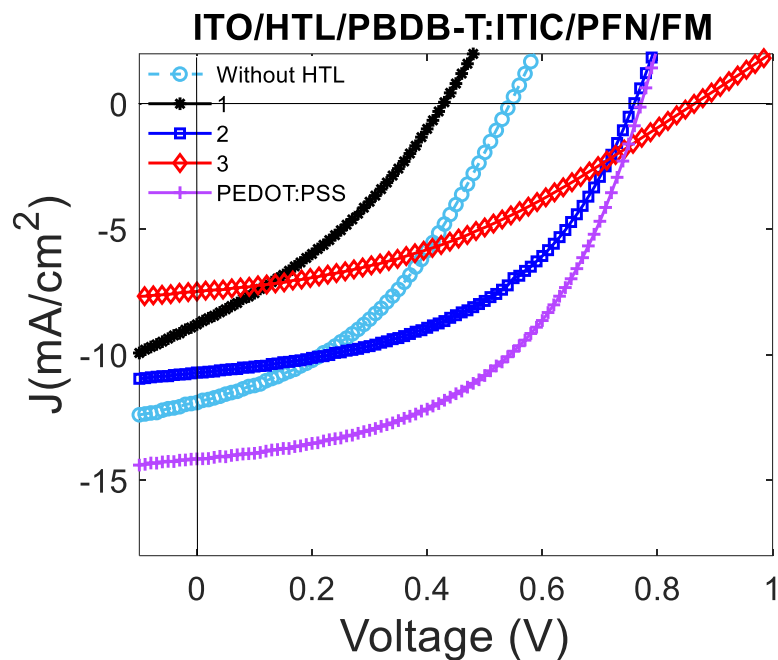


Fig. 7. J - V curve for PBDB-T:ITIC based OSC devices using the FM alternative top electrode vacuum-free deposited, with a HTL made of the C₁₈ ring *carbo*-mer **1**, **2** or **3**.

Table 3. PV parameters of OSCs with a HTL of the C₁₈ ring *carbo*-mer **1**, **2** or **3**. Most promising result for the dipolar *carbo*-benzene **2**.

Configuration: Glass/ITO/HTL/PBDB-T:ITIC/PFN/FM	V_{oc} (V)	J_{sc} (mA/cm ²)	FF	PCE (%)
Without HTL	0.55	12.1	0.41	2.7
1	0.44	9.0	0.33	1.3
2	0.76	10.7	0.48	3.9
3	0.86	7.5	0.38	2.5
PEDOT:PSS	0.77	14.2	0.50	5.5

Regarding previous reports on the use of SAMs as HTL in solar cells, Magomedov *et al.* [65], used a phosphonic acid-substituted carbazole derivative as hole transporting SAM allowing to

reach a PCE over 16% in regular configuration (PiN) perovskite-based solar cells (PSCs). Meanwhile, Yalcin *et al.* [17] reported a SAM of polyaromatic molecules functionalized with a carboxylic acid group leading to a PCE of 17.3%, which is among the highest efficiency reported to date for PiN PSCs by using an alternative HTL. Zhong *et al.* [66] observed a slight enhancement of the photovoltaic performance from 3.18% (with PEDOT:PSS) to 3.20% in OSCs having an ITO/HTL/CuI/P3HT:PC₆₁BM/ZnO/Al configuration with a SAM of 3-aminopropanoic acid as HTL. Very recently, Lin *et al.* [8] reached a PCE value of 18.03% upon using a SAM based on [2-(9*H*-carbazol-9-yl)ethyl]phosphonic acid (2PACz) along with a ternary active layer based on the very efficient PM6 donor. Most of the reported SAMs used as HTL are made of molecules bearing functional anchoring groups that interact with the metal oxide surface, and can influence the structure or thin film morphology and thus a suited conformation and contact of the deposited active layer [67].

4. Conclusions

In this work, self-assembled and conductive properties of 2D nano-films prepared from two *carbo*-benzenes and one *carbo*-quinoid were studied by SPM. STM images of the quadrupolar *carbo*-benzene **1** show that the molecular orientation could be controlled by the polarity of dimethylaniliny group influencing the dimensionality of assemblies on the surface, although no satisfactory STM images were obtained for the dipolar demethylated counterpart **2**. The STM images for the quadrupolar *carbo*-quinoid **3**, previously reported by our group, were related with those for **1**. Work function measured through KPFM resulted in values of $\phi = 5.60$ eV for **1**, 4.97 eV for **2**, and 4.82 eV for **3**. Hole mobility (μ) values extracted from local I-V plots by using c-AFM were $15 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **1**, $3 \times 10^{-6} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **2**, and $87 \times 10^{-8} \text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ for **3**. The high roughness (8-18 nm) and low homogeneity of thin films based on the selected *carbo*-mer derivatives, make them poorly adapted for use in the active layer of OSCs. Nevertheless, their ability to form SAMs, make them attractive for a possible use as alternative 2D HTL in OSCs, ensuring contact selectivity between the anode and the active layer. The highest PCE with a HTL of PEDOT:PSS was 5.5%, whereas a non-fully optimized value of 3.9% was obtained with a SA-HTM of the *carbo*-benzene **2**. Such C₁₈ ring *carbo*-mers could thus offer an alternative to conventional compounds used in the HTL, with the view to increasing the stability of OSCs.

Further improvements of both the *carbo*-mer structure by changing the nature and position of the substituents at the (pro)aromatic C₁₈ ring, and the preparation of the SAMs, should allow reaching higher OSC performance. Efforts will thus be pursued in this direction.

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Declaration of Competing Interest

The authors declare that they have no conflict of interest.

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