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Synthesis, characterization and crystal structures of two new platinum complexes with planar chiral 1,2-disubstituted ferrocenes containing phosphine and thioether donor groups¹

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¹ Dedicated to the memory of F. Albert Cotton, a caring mentor and a good friend.

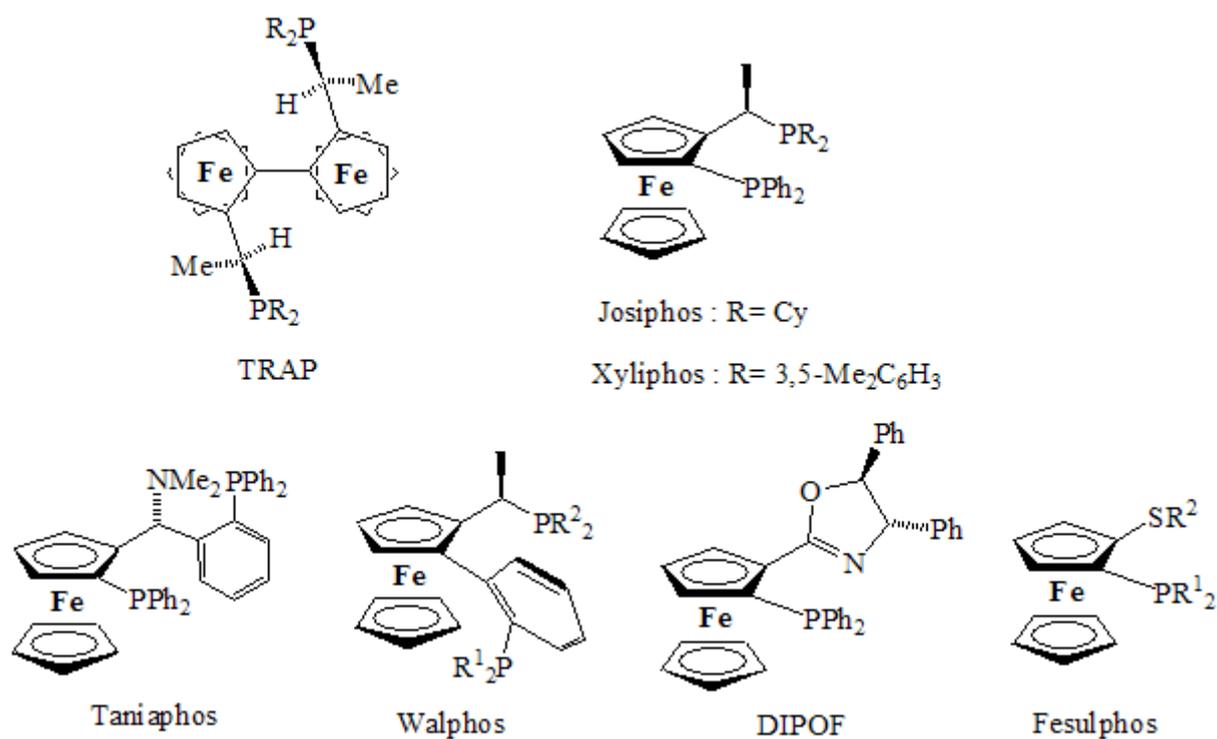
Abstract

The platinum complex [1,2-Fc(PPh₂)(CH₂SPh)]PtCl₂ **1** has been transformed into [1,2-Fc(PPh₂)(CH₂SPh)]PtCl(SnCl₃) **2** as a single isomer, in high yields, by reaction with SnCl₂. The SnCl₂ reagent inserts regioselectively in the Pt-Cl bond *trans* to the sulphur atom, while the sulphur atom configuration is fully controlled by the planar chirality of the P,S ligand. Complex **1** has also been efficiently reduced in the presence of diphenylacetylene to a single isomer of [1,2-Fc(PPh₂)(CH₂SPh)]Pt(PhCCPh), **5**, where the central chirality on sulphur is again maintained. In compound **5**, the central CC bond of the diphenylacetylene moiety acquires a strong double bond character. The structures of complexes **1** and **5** have been determined by X-ray diffraction on monocrystals. To the best of our knowledge, this study presents the first example of fully characterised L₂PtCl(SnCl₃) and L₂Pt(PhCCPh) complexes, L₂ being a P,S ligand.

Keywords: Platinum, ferrocene chemistry, phosphine ligands, thioether ligands, planar chirality

Introduction

Since its discovery by Pauson¹ and the first studies, which actively involved F. A. Cotton,² ferrocene has been used in fields of chemistry, like electrochemistry³, material science⁴, synthesis, and catalysis, particularly enantioselective catalysis because of their stability, easy introduction of planar chirality⁵ and special electronic and stereoproperties of the ferrocene skeleton. Amongst the chiral ferrocene-based ligands, enantiopure 1,2-disubstituted ferrocene derivatives, especially ferrocenyl-diphosphine ligands, played a major role.⁶ Important examples (see Scheme 1) are P,P ligands like TRAP⁷, Josiphos⁸, and particularly the industrially important Xyliphos⁹, Taniaphos¹⁰ and Walphos-type¹¹ ligands, P,N ligands,¹² especially phosphine-oxazolines like, for instance DIPOF¹³, and more recently P,S ligands like Fesulphos¹⁴.



Scheme 1

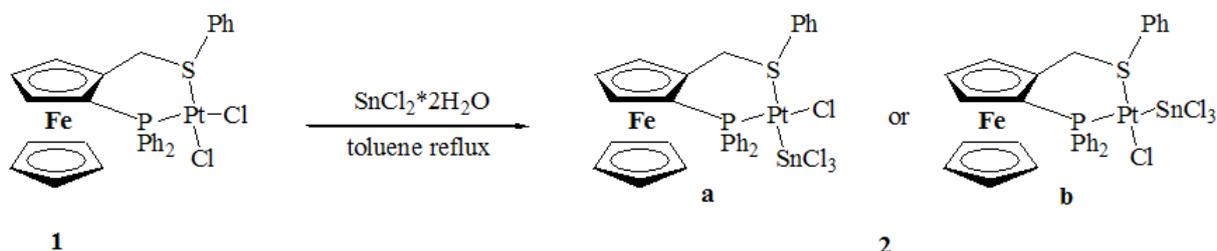
We have recently developed new chiral ferrocenyl (P,S) ligands with general formula $\text{CpFe}\{1,2\text{-C}_5\text{H}_3(\text{PPh}_2)(\text{CH}_2\text{SR})\}^{15}$, characterized by having only planar chirality, in both racemic and enantiomerically pure forms (R or S configuration)¹⁶ and briefly reported on their coordination chemistry¹⁷. These ligands could be successfully applied to a few asymmetric catalytic reactions, namely palladium-catalyzed allylic substitution^{15,18} and iridium-catalyzed ketone hydrogenation¹⁹. We have shown that the sulphur atom configuration is fully controlled by the planar chirality of the ligand.¹⁷ This is expected to be crucial in the transmission of chiral information to the catalytic metal. In this paper we report on the synthesis and characterization of two new platinum complexes containing the (P,S) ligand [1,2-Fc(PPh₂)(CH₂SPh)], which are potentially interesting for catalytic applications. One interesting aspect of this work is the stereochemical integrity shown by the sulfur donor atom while other metal-ligand bonds are implicated in different chemical transformations.

Results and Discussion

Several systems associating a platinum dichloride complex and tin dichloride have been used in the hydroformylation reaction of alkenes²⁰ and, particularly, in its asymmetric version.^{21,22} In these systems, platinum complexes bearing one chloro ligand and one trichlorotin ligand are considered to be important catalytic intermediates. Therefore, we decided to study the reaction of the platinum complex [1,2-Fc(PPh₂)(CH₂SPh)]PtCl₂, **1**, with SnCl₂.

The reaction of complex **1** with one equivalent of SnCl₂·H₂O in refluxing toluene yields a red solid, **2**, in good yields (90%) after one hour. This solid was first studied by NMR in CDCl₃. Only one signal at δ 1.21, different from that of the starting complex **1** (δ -0.8), was detected in the ³¹P NMR. This signal has two satellites corresponding to the Pt-P coupling ($J_{\text{Pt}} = 3517$ Hz) and two additional satellites corresponding to the Sn-P coupling ($J_{\text{Sn}} = 183$

Hz). This confirms the presence of tin in the product. In principle, two stereoisomers **a** and **b**, corresponding respectively to the insertion of SnCl_2 into the Pt-Cl bonds *trans* to S and *trans* to P, would be expected (see Scheme 2), but the ^{31}P NMR data show that only one compound is present. In addition, since the sulphur atom has central chirality, two different diastereoisomers for each geometric (*cis/trans*) isomer could exist in combination with the ferrocene planar chirality. The same phenomenon could take place for the starting material **1**, but only one diastereoisomer (with the sulphur substituent *anti* and the sulphur lone pair *syn* relative to the Fe atom) was found for this compound.^{17c} Thus, the formation of only one product from the single diastereoisomer of the starting compound shows that the reaction is not only regioselective with respect to the reactive Pt-Cl bond but also diastereoselective with respect to the sulphur configuration (either 100% retention or 100% inversion).



Scheme 2

For the similar complexes *cis*-(PR_3)₂PtCl(SnCl_3), it has been shown that the Pt-P coupling constants are greater for the P atom *trans* to Cl (3090-4910 Hz) than for the P atom *trans* to SnCl_3 (2520-4110 Hz).^{22,23} For compound **2**, however, the Pt-P coupling constant ($J_{\text{PPt}} = 3517$ Hz) is inside both ranges and does not allow an unambiguous structural assignment. On the other hand, the Sn-P coupling constant is more diagnostic because its value for *cis*-(PR_3)₂PtCl(SnCl_3) is very different when P atom is *trans* to Cl (78-235 Hz) or *trans* to SnCl_3 (373-5600 Hz).^{22,23} When compared with these literature precedents, the value

for **2** ($J_{\text{PSn}} = 183$ Hz) strongly suggests that the phosphorus atom is in *trans* of the chlorine atom, namely that isomer **a** is the reaction product.

The ^{195}Pt chemical shift²⁴ for compound **2** (δ -4273) is, as expected, slightly downfield from those observed for the corresponding complexes bearing two phosphine ligands (δ -4710 to -4910)^{24a,25,22c} as already observed within the *cis*-PtCl₂L₂ family.^{Erreur ! Signet non défini.} The chemical shift of **2** is upfield relative to the corresponding dichloro complex **1** ($\delta = -4093$ ppm), as found in similar pairs.^{24a} The ^{13}C NMR and ^1H NMR data also confirm that only one compound is present in solution. To the best of our knowledge, this study is the first characterization of a L₂PtClSnCl₃ complex, with L₂ being a P,S ligand.

Single crystals of **2** were obtained by slow diffusion of pentane in a dichloromethane solution. The structure obtained by X ray diffraction (Figure 1) shows a square planar geometry around the platinum atom with the SnCl₃ group in *trans* to the sulphur atom. The (**R/S**)-**2** molecule is to the best of our knowledge the first (phosphine-thioether)-PtCl(SnCl₃) complex characterized by X-ray analysis. The asymmetric unit contains two nearly identical molecules with the same ($R_{\text{Fc}}, R_{\text{S}}$) or ($S_{\text{Fc}}, S_{\text{S}}$) configuration. The phenyl substituent on the sulphur atom is oriented *anti* relative to the Fe atom, *i.e.* the observed diastereomer is the same as that of the starting compound, meaning that the the sulphur configuration of the (**R/S**)-**1** precursor is preserved during the reaction with SnCl₂. The PtPSClSn fragment of each independent molecule shows the expected planar coordination geometry around the central Pt atom as evidenced from: (i) the sum of the four intra- and interligand cis angles, found equal to 360.5°(359.9°), and (ii) from the angle between the normal to the two planes defined by the PtPS and PtClSn fragments which is equal to 3.38(4)° (7.9(1)°) thus deviating only a little from the ideal value of 0° required for an undistorted planar geometry. The geometry within the square plane compares well with related structures containing the PtCl(SnCl₃) fragment as shown in Table 1. It is worth noting that the Pt-Sn bond is

significantly shorter in (**R/S**)-**2** than in complexes **3**^{22c} and **4**^{23b} where the SnCl₃ group is *trans* with respect to a phosphorus atom. This result was expected, because the *trans* effect of the thioether group is weaker than that of a phosphine. The angles around the platinum atom are similar for the three complexes.

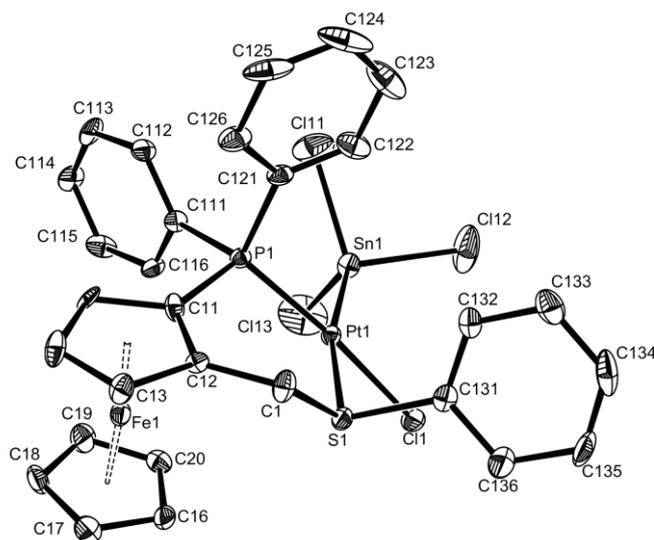
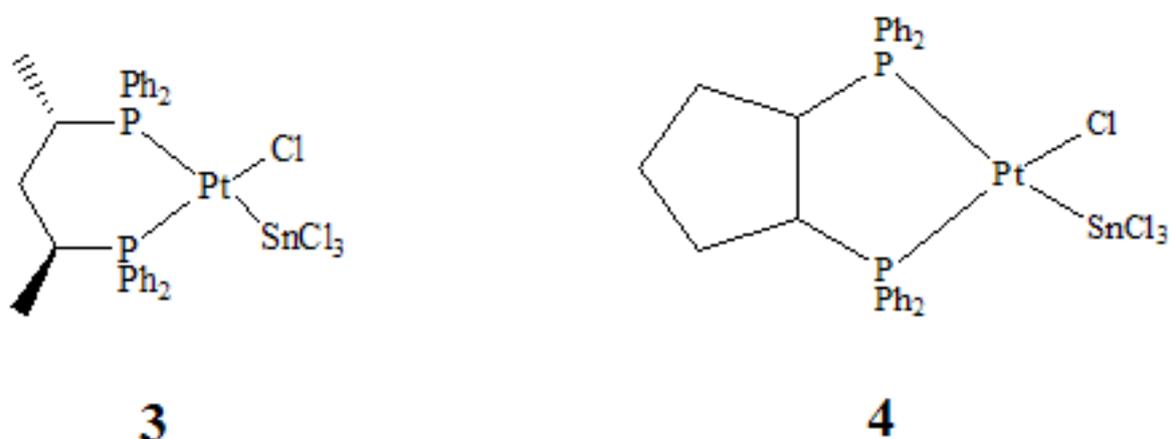


Figure 1. Molecular view of one of the two crystallographically independent molecules of (**R/S**)-**2** with the atom-labelling scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms have been omitted for clarity.



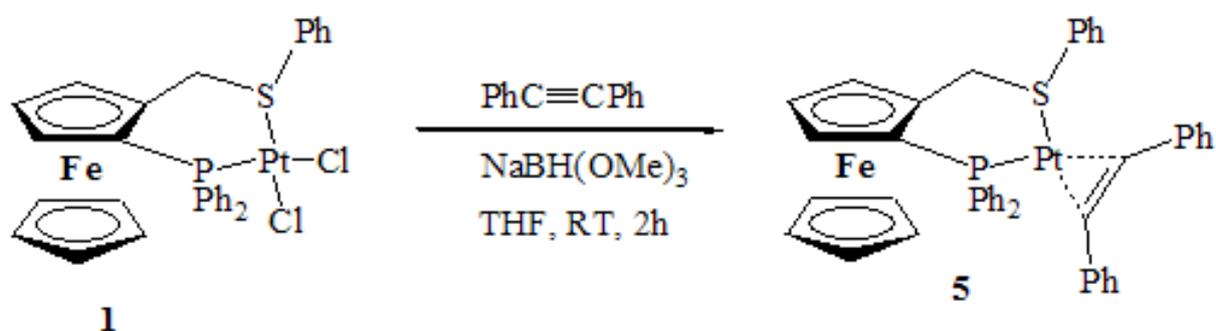
Scheme 3

Table 1. Comparison of bond distances (Å) and angles (deg) for complexes **2**,^a **3** and **4**

	2		3	4
Pt – P (<i>trans</i> to Cl)	2.223(2)	2.235(2)	2.234	2.275(5)
Pt – X (<i>trans</i> to Sn) ^b	2.357(2)	2.347(2)	2.309	2.281(5)
Pt - Sn	2.5331(6)	2.5312(6)	2.582	2.572(2)
Pt - Cl	2.350(2)	2.334(2)	2.347	2.283(9)
P – Pt – X ^b	95.85(7)	94.88(7)	93.83	87.0(2)
Cl – Pt - Sn	83.98(5)	82.71(5)	81.55	85.2(3)
Reference	This work		22c	23b

^aThe parameters for the two independent molecules are listed in separate columns. ^bX = S for **2**; P for **3** and **4**.

In order to obtain Pt⁰ complexes, which are potential catalyst for several reactions^{26,27}, the reactivity of complex **1** was tested with diphenylacetylene in presence of sodium tris(methoxy)borohydride as a reducing agent (see scheme 4), following a method described by Glueck et al.²⁷ for the synthesis of a *trans*-stilbene Pt(0) complex (scheme 4). The expected reactivity is the reduction of platinum(II) to Pt⁰, the coordination sphere being completed by diphenylacetylene acting as L ligand. The reaction was carried out in THF room temperature for two hours. A brown compound, [1,2-Fc(PPh₂)(CH₂SPh)]Pt(PhC≡CPh) (**5**), was obtained in 89% yield.



Scheme 4

The ^{31}P NMR spectrum of **5** shows one single singlet at δ 12.6 with two satellites corresponding to the platinum-phosphorus coupling ($J_{\text{Pt}} = 3758$ Hz) different from that of the starting complex **1** ($\delta = -0.8$ ppm, $J_{\text{Pt}} = 3608$ Hz) indicating the presence of only one compound. Similarly, the ^1H NMR spectrum presents only one singlet for a 1,2-disubstituted ferrocene and also a more complex signal corresponding to 25 aromatic protons (15 from the P,S ligand and 10 from diphenylacetylene). The structure proposed in Scheme 4 is also supported by the ^{13}C NMR data, which clearly show the presence of a dissymmetrically coordinated diphenylacetylene. In particular, the signals corresponding to the 2 carbons of the $\text{C}\equiv\text{C}$ bond are found at 127.1 ppm and 126.5 ppm.

Single crystals of (**R/S**)-**5** were obtained by slow diffusion of pentane in a dichloromethane solution. The structure obtained by X ray diffraction is presented in Figure 2. Because the synthesis of this complex was made starting from a racemic mixture of **1**, both enantiomers of **5** are present in the crystal. The asymmetric unit of compound (**R/S**)-**5**, like that of compound **2** discussed above, contains two crystallographically independent molecules, both containing the same relative configuration for the two elements of chirality, ($R_{\text{Fc}}, R_{\text{S}}$) or ($S_{\text{Fc}}, S_{\text{S}}$). Once again, the chemical transformation from **1** to **5** does not change the initial configuration of the sulphur atom, because the Ph substituent on the S atom is oriented *anti* to the Fe atom in both compounds. The X ray structure shows that the phenyl acetylene is η^2 bonded to the platinum, as proposed by the NMR data. Thus, **5** is a platinum(0) ferrocenylphosphine-thioether complex. To the best of our knowledge, no other phosphine-thioether Pt^0 complex has previously been described in the literature. We can compare the structural data of complex **5** with those of bis(phosphine) Pt^0 complexes containing diphenylacetylene, **6**²⁸, **9**²⁹, **10-12**³⁰ and 1,4-diphenylbutadiyne, **7**³¹, **8**³² (scheme 5). The principal structural data for these complexes are collected in Table 2. The C1-C2 distance

corresponds to the C-C bond that is η^2 bonded to the platinum atom. Positions α and β correspond to the atoms directly bonded to C1 and C2.

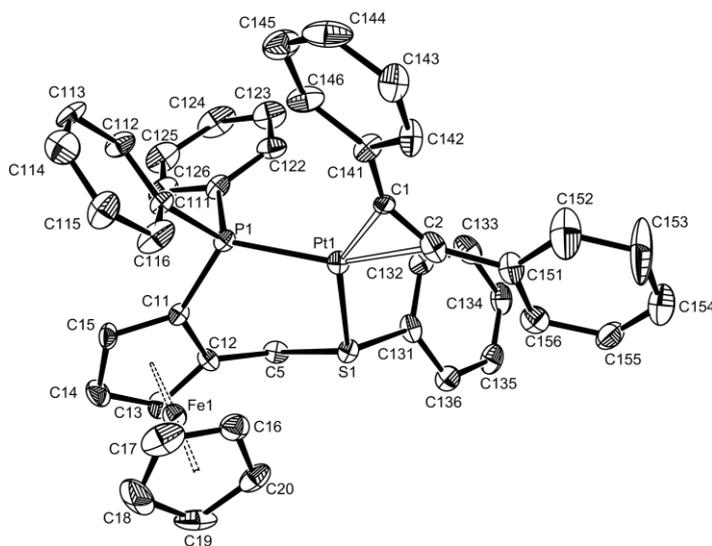
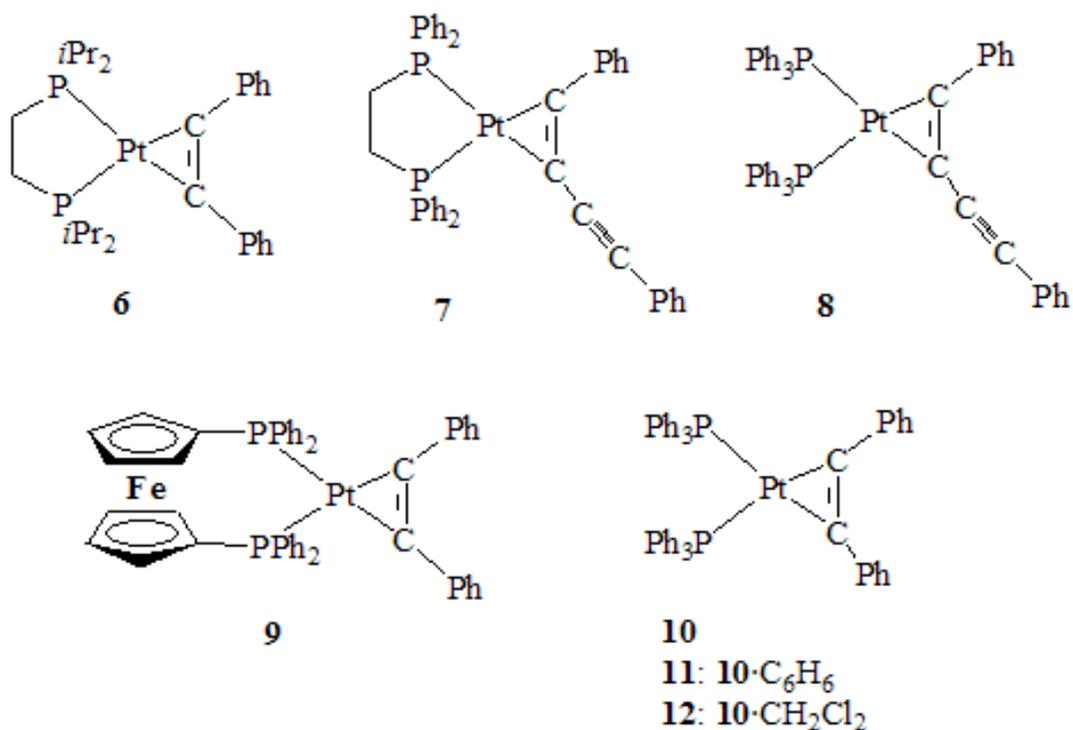


Figure 2. Molecular view of one of the two crystallographically independent molecules of (*R/S*)-**5** with the atom-labelling scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms have been omitted for clarity.



Scheme 5

Table 2. Comparison of bond distances (Å) and angles (deg) for complexes **5** and **6-12**.^a

	5		6	7	8	9	10	11	12	
Pt - P	2.263(3)	2.260(3)	2.2577(9)	2.25(1)	2.273(2)	2.2714(16)	2.2885(7)	2.2820(9)	2.2903(7)	2.2895(7)
Pt - X^b	2.351(3)	2.340(3)	2.2672(10)	2.26(1)	2.279(2)	2.2679(18)	2.2895(6)	2.2910(8)	2.2888(7)	2.2812(7)
Pt - C1	2.007(12)	2.042(11)	2.029(3)	1.98(3)	2.036(8)	2.050(7)	2.057(2)	2.047(3)	2.057(3)	2.047(3)
Pt - C2	2.039(11)	2.073(11)	2.047(3)	2.02(3)	2.041(7)	2.040(7)	2.049(2)	2.048(3)	2.050(3)	2.048(3)
C1 - C2	1.278(15)	1.290(14)	1.301(5)	1.39(4)	1.305(11)	1.293(9)	1.295(4)	1.280(5)	1.289(4)	1.280(4)
C1 - Pt - C2	36.8(4)	36.5(4)	37.24(13)	39.0(11)	37.3(3)	36.9(2)	36.77(10)	36.43(13)	36.58(11)	36.43(11)
Cα - C1 - C2	144.3(11)	144.8(11)	144.4(3)	138.9(7)	140.7(7)	138.6(7)	141.8(2)	142.3(3)	138.6(3)	139.5(3)
C1 - C2 - Cβ	146.6(12)	145.2(12)	140.7(3)	142.8(4)	146.7(8)	147.0(7)	140.7(2)	142.6(3)	143.3(3)	145.2(3)
Ref.	This work		28	31	32	29	30	30	30	

^aFor compounds **5** and **12**, the parameters for the two independent molecules are listed in separate columns. ^bX = S for **5** ; P for **6-12**.

The structural analysis reveals the expected distorted square-planar geometry at the metal centre with the alkyne ligand in the P-Pt-S plane, the largest deviation from the PPtSC1C2 mean plane being 0.038(6) Å at C1 [0.039(6) at C3]. The elongated C1-C2 bond, 1.278(15) Å [1.290(14)] and the bending of the C≡C-C angles [144(1)° to 147(1)°] reflect the major contribution of π^* back-donation. This bond length is quite close to those observed in the related complexes with two phosphine donors (Table 2).

Conclusion

We have presented here the first synthesis and characterization of platinum complexes of type $L_2PtCl(SnCl_3)$ and $L_2Pt(PhCCPh)$ where L_2 is a P,S ligand. The study has shown that the configuration at sulphur is preserved on going from **1** to **2** and **5**. It is worth noting that the synthesis of complex **1** from the corresponding ferrocene phosphine-thioether ligand and $PtCl_2(CH_3CN)_2$ is also stereoselective, yielding only one of the two possible diastereoisomers. The same occurs for all other analogous 1,2-Fc(PPh₂)(CH₂SR) ligands with R = Ph, Et, tBu, the observed product having in all cases the sulphur atom lone pair oriented toward the unsubstituted Cp ring. Therefore, the planar chirality of the sulphur atom controls the configuration at the sulphur atom.

Acknowledgement

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Experimental part

All reactions were carried out under dry argon by using Schlenk glassware and vacuum line techniques. Solvents were freshly distilled from standard drying agents. ^1H , $^{13}\text{C}\{^1\text{H}$, $^{31}\text{P}\}$, $^{119}\text{Sn}\{^1\text{H}\}$, $^{31}\text{P}\{^1\text{H}\}$, $^{195}\text{Pt}\{^1\text{H}\}$ NMR spectra were recorded on a Bruker WMX 400 instrument operating at 400, 162, 186.6, 100, and 86 MHz respectively. Chemical shifts are reported in parts per million (ppm) relative to Me_4Si (^1H and ^{13}C), Me_4Sn (^{119}Sn), 85% H_3PO_4 (^{31}P) or aqueous solution of Na_2PtCl_6 (^{195}Pt). Mass spectra were obtained on a Nermag R10-10 instrument.

Synthesis of (R/S)-(2-diphenylphosphine(phenylthiomethyl)ferrocene)chloro-(trichlorostannyl)platinum(II), (R/S)-2.

In a Schlenk tube under argon, 82 mg (0.1 mmol) of complex **(R/S)-1** were dissolved in 5 mL of distilled toluene. $\text{SnCl}_2 \cdot 2\text{H}_2\text{O}$ (23 mg, 0.1 mmol) were then added. The solution was refluxed for one hour, yielding a red solid in a yellow solution. The solid was filtered, washed with pentane and dried under vacuum. Yield: 93 mg, 90%. ^1H NMR (500 MHz, CD_2Cl_2): δ 7.69-7.56 (12H, m, Ph); 7.53-7.49 (1H, m, Ph); 7.45-7.41 (2H, m, Ph); 4.75 (1H, pseudo q, $J_{\text{HH}} = J_{\text{HP}} = 1.6$ Hz, subst Cp); 4.71 (5H, s, Cp); 4.53 (1H, dt, $J_{\text{HH}} = 1.5$ Hz, $J_{\text{HP}} = 1$ Hz, subst Cp); 4.47 (1H, d(AB), $J_{\text{HH}} = 13$ Hz, $\underline{\text{CH}}_2\text{-Cp}$); 3.69 (1H, pseudo q, $J_{\text{HH}} = J_{\text{HP}} = 1.6$ Hz, subst Cp); 3.63 (1H, d(AB), $J_{\text{HH}} = 13$ Hz, $\underline{\text{CH}}_2\text{-Cp}$). $^{13}\text{C}\{^1\text{H}\}$ NMR (500 MHz, CDCl_3): δ 133.9 (d, $J_{\text{CP}} = 12$ Hz, PPh_2); 133.2 (s, SPh); 132.9 (d, $J_{\text{CP}} = 12$ Hz, PPh_2); 132.8 (d, $J_{\text{CP}} = 3$ Hz, PPh_2); 132.7 (d, $J_{\text{CP}} = 3$ Hz, PPh_2); 131.2 (s, SPh); 129.7 (d, $J_{\text{CP}} = 72$ Hz, quat PPh_2); 129.7 (s, SPh); 129.6 (d, $J_{\text{CP}} = 13$ Hz, PPh_2); 129.2 (d, $J_{\text{CP}} = 12$ Hz, PPh_2); 127.8 (s, SPh); 127.0 (d, $J_{\text{CP}} = 64$ Hz, quat PPh_2); 83.5 (d, $J_{\text{CP}} = 14$ Hz, quat Cp); 76.1 (d, $J_{\text{CP}} = 7$ Hz, subst Cp); 74.7 (d, $J_{\text{CP}} = 7$ Hz, subst Cp); 71.5 (s, Cp); 69.3 (d, $J_{\text{CP}} = 8$ Hz, subst Cp); 65.3 (d, $J_{\text{CP}} = 73$ Hz, quat Cp); 39.5 (s, $\underline{\text{CH}}_2\text{-Cp}$). ^{31}P NMR (500 MHz, CDCl_3): δ 1.21 ($J_{\text{PPt}} = 3517$ Hz, J_{PSn}

= 183 Hz). ^{195}Pt NMR (500 MHz, CDCl_3): δ -4273. ^{119}Sn NMR (500 MHz, CDCl_3): δ -263.4. MS (DCI; NH_3): m/z = 966 ($\text{M} + \text{NH}_3$, 8 %); 776 ($\text{M} - \text{SnCl}_2 + \text{NH}_4$, 51%).

Synthesis of (R/S)-(2-diphenylphosphino-(phenylthiomethyl)ferrocene)(diphenylacetylene)platinum(0), (R/S)-5. In a Schlenk tube under argon, 80 mg (0.1 mmol) of complex (R/S)-1 and 18 mg (0.1 mmol) of diphenylacetylene were dissolved in 5 mL of dry THF. A solution of $\text{NaBH}(\text{OMe})_3$ (78 mg, 0.6 mmol) in 3 ml of dry THF was then added. The solution was stirred for two hours at room temperature. A small amount of white solid that had formed in the brown solution was eliminated by filtration. The solvent was then evaporated to dryness yielding the product as a red-brown solid. The solid was washed with pentane and dried under vacuum to yield a red solid (77 mg, 89% yield). ^1H NMR (500 MHz, CDCl_3): δ 7.91-7.85 (3H, m, Ph); 7.73-7.63 (2H, m, Ph); 7.57-7.46 (3H, m, Ph); 7.42-7.21 (12H, m, Ph); 7.14-6.98 (5H, m, Ph); 4.50 (1H, pseudo q, $J_{\text{HH}} = J_{\text{HP}} = 2$ Hz, subst Cp); 4.44 (1H, d(AB), $J_{\text{HH}} = 12$ Hz, $\text{CH}_2\text{-Cp}$); 4.38 (5H, s, Cp); 4.31 (1H, t, $J_{\text{HH}} = 2$ Hz, subst Cp); 3.93 (1H, d, $J_{\text{HH}} = 2$ Hz, subst Cp); 3.62 (1H, d(AB), $J_{\text{HH}} = 12$ Hz, $\text{CH}_2\text{-Cp}$). $^{13}\text{C}\{^1\text{H}\}$ NMR (500 MHz, CDCl_3): δ 137.0 (d, $J_{\text{CP}} = 4$ Hz, quat SPh); 136.5 (d, $J_{\text{CP}} = 42$ Hz, quat PPh₂); 134.3 (d, $J_{\text{CP}} = 15$ Hz, o PPh₂); 133.8 (d, $J_{\text{CP}} = 47$ Hz, quat PPh₂); 133.4 (d, $J_{\text{CP}} = 15$ Hz, o PPh₂); 132.9 (s, SPh); 132.6 (s, SPh); 129.9 (s, $\text{PhC}\equiv\text{CPh}$); 129.8 (d, $J_{\text{CP}} = 2.5$ Hz, p PPh₂); 129.3 (d, $J_{\text{CP}} = 2.5$ Hz, p PPh₂); 129.2 (s, SPh); 128.9 (s, quat $\text{PhC}\equiv\text{CPh}$); 128.6 (s, $\text{PhC}\equiv\text{CPh}$); 128.4 (s, quat $\text{PhC}\equiv\text{CPh}$); 128.3 (d, $J_{\text{CP}} = 10$ Hz, m PPh₂); 128.2 (s, SPh); 127.7 (s, $\text{PhC}\equiv\text{CPh}$); 127.7 (s, $\text{PhC}\equiv\text{CPh}$); 127.6 (d, $J_{\text{CP}} = 10$ Hz, m PPh₂); 127.1 (s, $\text{PhC}\equiv\text{CPh}$); 126.5 (s, $\text{PhC}\equiv\text{CPh}$); 125.8 (s, $\text{PhC}\equiv\text{CPh}$); 125.4 (s, $\text{PhC}\equiv\text{CPh}$); 87.4 (d, $J_{\text{CP}} = 18$ Hz, quat Cp); 74.6 (1C, d, $J_{\text{CP}} = 6$ Hz, subst Cp); 73.4 (d, $J_{\text{CP}} = 28$ Hz, quat Cp); 73.2 (s, subst Cp); 70.7 (s, Cp); 66.8 (1C, d, $J_{\text{CP}} = 5$ Hz, subst Cp); 42.3 (d, $J_{\text{CP}} = 7$ Hz, $\text{CH}_2\text{-Cp}$). ^{31}P NMR (500 MHz, CDCl_3): δ 12.6 ($J_{\text{Ppt}} = 3758$ Hz).

X-ray crystallography. A single crystal of each compound was mounted under inert perfluoropolyether at the tip of a glass fibre and cooled in the cryostream of an Oxford-Diffraction XCALIBUR CCD diffractometer. Data were collected using the monochromatic MoK α radiation ($\lambda = 0.71073$). The structures were solved by direct methods (SIR97)³³ and refined by least-squares procedures on F^2 using SHELXL-97.³⁴ All H atoms attached to carbon were introduced in idealised positions and treated as riding on their parent atoms in the calculations. The drawing of the molecules was realised with the help of ORTEP3.³⁵ Crystal data and refinement parameters are shown in Table 3. Bond distances and angles are given in Tables 1 and 2.

Crystallographic data (excluding structure factors) have been deposited with the Cambridge Crystallographic Data Centre as supplementary publication no. CCDC 675787 & 675788. Copies of the data can be obtained free of charge on application to the Director, CCDC, 12 Union Road, Cambridge CB2 1EZ, UK (fax: (+44) 1223-336-033; e-mail: deposit@ccdc.cam.ac.uk).

Table 3. Crystal data and structure refinement

Identification code	(R/S)-2	(R/S)-5
Empirical formula	C ₅₉ H ₅₂ Cl ₁₀ Fe ₂ P ₂ Pt ₂ S ₂ Sn ₂	C ₄₃ H ₃₅ Cl ₁₀ FePtS
Formula weight	990.41	865.68
Temperature, K	180(2)	180(2) K
Crystal system	Monoclinic	Triclinic
Space group	P 2 ₁ /c	P -1
a, Å	10.4191(5)	12.9671(19)
b, Å	17.8962(9)	16.1949(19)
c, Å	35.1630(17)	18.672(2)
α , °	90.0	96.936(9)
β , °	93.004(4)	103.775(11)
γ , °	90.0	100.870(11)
V, Å ³	6547.6(6)	3683.2(8)

Z	8	4
D _{calc} , Mg/m ³	2.009	1.561
μ, mm ⁻¹	5.998	4.317
F(000)	3784	1712
Crystal size, mm ³	0.27x0.041x0.039	0.33x0.16x0.15
θ °, range	3.00 to 25.68	2.82 to 26.37
Reflections collected	44982	28106
Unique reflections [R _{int}]	12418 [0.0796]	15016 [0.0865]
Obsd reflections [I > 2σ(I)]	7133	7907
Completeness	99.8	99.7
Absorption correction	Multi-scan	Multi-scan
T _{min} /T _{max}	0.3745 / 0.7984	0.1588 / 0.3062
parameters	712	869
GOF on F ²	0.814	0.901
R ₁ , wR ₂ [I > 2σ(I)]	0.0440, 0.0445	0.0634, 0.1442
R ₁ , wR ₂ (all data)	0.0971, 0.0527	0.1034, 0.1618
Δρ _{min} , Δρ _{max} (e Å ⁻³)	-0.999, 1.310	-2.581, 2.817

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