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Towards redox-switchable organocatalysts based on bidentate halogen bond donors†

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Redox-active bidentate halogen bond donors based on halopyridinium groups as halogen-bond donating units were synthesized and their structures were elucidated by X-ray diffraction analyses and DFT calculations. *Via* reversible twofold reduction, these dicationic species can be transformed to neutral compounds which should be much weaker Lewis acids. The corresponding electrochemical data were obtained, and CV as well as UV-vis and NMR techniques were also used to determine binding constants of these halogen bond donors to halides. While all titrations agree on the relative order of binding strengths (with chloride being bound strongest), there are marked deviations in the overall affinity constants which are discussed. In contrast to earlier azo-bridge analogues, the ethylene-linked variants presented herein do not oxidize halides, and thus the novel halogen bond donors could also be used as Lewis acidic organocatalysts in a halide abstraction benchmark reaction, yielding a performance similar to bis(haloimidazolium)-derived catalysts.

Introduction

Recently we have demonstrated the possibility to tune and control the strength of halogen bonding (XB) by electrochemically switching the redox state of the respective electroactive XB donors or acceptors. The observed shift of the standard potentials can be attributed to a stabilization of the oxidized or reduced redox probes *via* charge-assisted halogen bonding. The first proof of concept was brought for *p*-quinones playing the role of electrochemically active XB-acceptors.¹ Since then it has been applied to various types of redox active XB-donors such as iodinated TTF,² ferrocene³ or viologen⁴ derivatives confirming the interesting potential of electrochemical techniques as a powerful tool to detect and control XB in solution.⁵ In parallel we described the use of halogenated azo-bis-pyridinium derivatives as activators in a proof-of-principle reaction (Scheme 1, X₄-2 derivatives).⁶ Even though the activity of these halogen bond donors was very promising, the analysis of the reaction was severely complicated by a side reaction, in which the bromide liberated from the S_N1 substrate was oxidized by the

halogen bond donor to form elemental bromine and the hydrazine form of the donor. Careful comparison experiments could still demonstrate that the azo-bis-pyridinium derivatives X₄-2 (Scheme 1) were active through halogen bonding, but the observed redox activity does of course strongly limit the potential application of these halogen-based Lewis acids in halide abstraction reactions. The situation should be different, however, for the C analogous tetrachloro, tetrabromo and tetraiodo stilbene derivatives X₄-1 (Scheme 1), which should be much less sensitive to reduction. In fact, purposeful reduction of these compounds by electrochemistry would allow to basically switching off their halogen bonding capabilities. Changing the XB-donor strength of such bidentate ligands *via* a reversible electrochemical reaction can lead to a redox-switchable receptor. In this work, we report the preparation of the novel receptors X₄-1, the investigation of their XB-donor properties depending on the oxidation state and the evaluation of their catalytic properties. Their interesting potential as redox-switchable catalysts^{7,8} in organocatalysis is discussed.

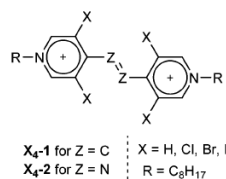
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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0cp06612e

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Scheme 1 Stilbene and azo derivatives X₄-1 and X₄-2.

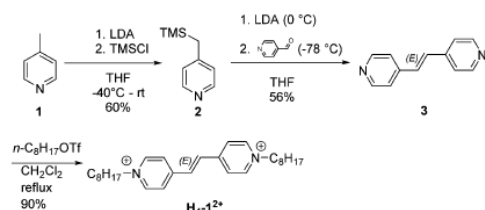
In-depth studies on the determination of the thermodynamic binding constants towards anions have been performed using multiple analytical techniques affording valuable quantitative data on XB strength in solution, a domain in which systematic data are still longed for.⁹

Results and discussion

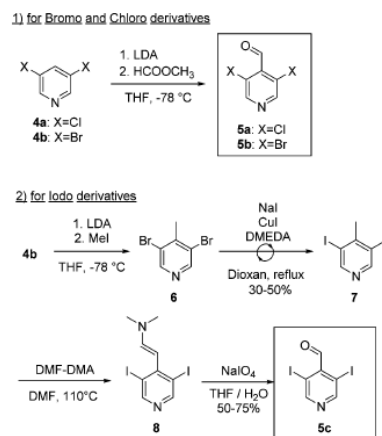
1. Synthesis and characterization

Different synthetic strategies have been undertaken for the preparation of the target compounds X_4-1 ($X = H, Cl, Br, I$) (Scheme S1, ESI[†]). In order to access the target compounds we followed two separate pathways, starting either from 4-picolin (1) or 3,5-dihalopyridines (4a–b). For the non-halogenated species 3 (Scheme 2), 4-picolin (1) was transformed into 4-picolinyl-trimethylsilane (2) by lithiation with LDA at low temperatures and subsequent quenching with TMSCl. In a following Peterson olefination, 2 was lithiated with LDA and treated with isonicotinaldehyde.¹⁰ After several successive precipitations diastereomeric 3 was obtained as the pure *E*-isomer. Alkylation with octyltriflate afforded H_4-1^{2+} , following a published procedure.⁶ For the synthesis of the halogenated compounds two distinct procedures were used, depending on the nature of the halogen atom (Scheme 3). Formylation of the 3,5-dichloro and 3,5-dibromo pyridines 4a–4b afforded the respective aldehydes 5a–5b.¹¹ The analogue 3,5-diiodoisonicotinaldehyde (5c) could not be obtained¹² with this procedure, probably due to the buttressing effect of the flanking iodines.¹³ After methylation of 4b with CH_3I and halide exchange on the resulting 3,5-dibromo-4-picolin (6), the diiodinated derivative 7 was obtained. Refluxing of 7 in DMF–DMA¹⁴ afforded enamine 8 which could be oxidized with sodium periodate¹⁵ to the key intermediate 5c.

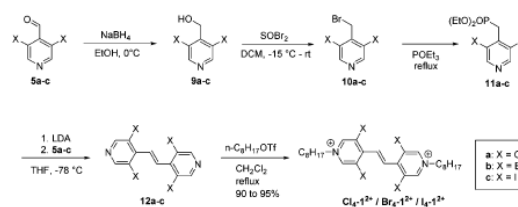
The aldehydes 5a–c were reduced with $NaBH_4$ affording the corresponding carbinols 9a–c (Scheme 4). Conversion to the picolinyl bromides 10a–c was conducted with thionyl bromide in methylene chloride. An Arbuzov reaction of picolinyl bromides with triethylphosphite in a sealed vessel at 110 °C yielded quantitatively the corresponding diethyl phosphonates 11a–c. After lithiation with LDA and subsequent addition of the respective aldehydes 5a–c, the corresponding azastilbenes 12a–c were isolated after filtration and trituration. Alkylation with *n*-octyltriflate afforded the vinylogous bipyridinium derivatives Cl_4-1^{2+} , Br_4-1^{2+} and I_4-1^{2+} .



Scheme 2 Synthesis of the non-halogenated derivative H_4-1^{2+} .



Scheme 3 Synthesis of the 3,5-dihalogeno-isonicotinaldehyde 5a, 5b and 5c.



Scheme 4 Synthesis of the targeted molecules Cl_4-1^{2+} , Br_4-1^{2+} and I_4-1^{2+} .

The electrochemical behaviour of the three extended viologens H_4-1^{2+} , Br_4-1^{2+} and I_4-1^{2+} was investigated by cyclic voltammetry (CV) in DMF with tetrabutylammonium hexafluorophosphate (nBu_4PF_6) as supporting electrolyte. Two different behaviors were observed. The CV of H_4-1^{2+} exhibits two reversible single-electron reduction waves (Fig. 1, black trace) with formal standard potentials (vs. SCE) of -0.46 and -0.66 V corresponding to the sequential

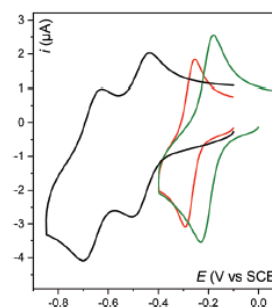


Fig. 1 CVs of 0.1 mM of I_4-1^{2+} (red trace), Br_4-1^{2+} (green trace) and H_4-1^{2+} (black trace) in 0.1 M nBu_4PF_6 /DMF on a glassy carbon electrode (3 mm-diameter) at 298 K. $\nu = 0.1$ V s^{-1} .

formation of the radical cation ($\text{H}_4\text{-1}^{\bullet+}$) and the neutral fully reduced species ($\text{H}_4\text{-1}$), respectively (Table S1, ESI[†]). In contrast, CVs recorded for $\text{Br}_4\text{-1}^{2+}$ and $\text{I}_4\text{-1}^{2+}$, present only one single reversible wave (Fig. 1, green and red traces) corresponding to a two-electron transfer (ET) process respectively at -0.20 and -0.27 V vs. SCE (Table S1, ESI[†]) leading directly to the neutral species. Owing to the electron withdrawing effect of the halo substituents, the formal standard potentials of the halogenated derivatives are significantly less negative compared to the parent compound $\text{H}_4\text{-1}^{2+}$. The different mechanism (two single-electron vs. one two-electron transfer process) can be attributed to steric hindrance.

The presence of four bromo or iodo substituents probably causes a strong conformational torsion leading to a disruption of the π -conjugated system of the molecules. In the case of $\text{I}_4\text{-1}^{2+}$, an additional wave was observed at -1.70 V, corresponding to the signature of a dissociative C-I bond reduction (Fig. S6 red trace, ESI[†]).¹⁶ In all cases, a linear dependence of the peak currents on the square root of the scan rate confirms diffusional electrochemical reactions without any adsorption phenomena of the studied compounds (Fig. S7-S9, ESI[†]).

Crystals of $\text{I}_4\text{-1}^{2+}$ triflate salt [$\text{I}_4\text{-1}^{2+}(\text{TfO}^-)_2$] were obtained by slow evaporation of a solution in an ACN/DMSO/benzene solvent mixture, the chloride complex [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)(\text{TfO}^-)$] was crystallized from a DMF solution, and the dichloride [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)_2$] by the slow solvolysis of benzhydryl chloride in acetonitrile and evaporation of the solvent over months, since the direct addition of chloride salts only lead to precipitation (Fig. 2 and Fig. S10-S12, ESI[†]).

The three compounds crystallized in the triclinic space group $P\bar{1}$. $\text{I}_4\text{-1}^{2+}$ and its dichloride contain only one bipyridine molecule in the unit cell, displaying an inversion center on the ethylene bridge (Fig. 2 and Table 1). The triflate anion in [$\text{I}_4\text{-1}^{2+}(\text{TfO}^-)_2$] is bound by I1 (2.94 Å, 84% vdW) while I2 interacts rather loosely with O1 and O2. The distances are about 3.44 Å, which is only 1% shorter than the sum of the vdW radii. The C-I2...O angles are 147° and 164.5° , so that the interaction cannot be considered a XB but a mere packing effect. In contrast, the crystal structure of the complex [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)(\text{TfO}^-)$] displays four non covalent I...Cl contacts ($d_{\text{I}\cdots\text{Cl}} = 3.09\text{--}3.31$ Å) of 11%

Table 1 Halogen bond geometry in the crystal structures of compounds $\text{I}_4\text{-1}^{2+}$ and [$\text{I}_4\text{-1}^{2+}, \text{Cl}^-$]: inter atomic distances (d) and angles (\angle)

$[\text{I}_4\text{-1}^{2+}(\text{TfO}^-)_2]^a$	$d_{\text{I}\cdots\text{O}}^a$	% vdW	$\angle \text{X-I}\cdots\text{O}^b$	$d_{\text{C}\cdots\text{I}}^a$
C2-I1...O1	2.94	84.0	173.5	2.09
C4-I2...O1	3.45	98.6	147.0	2.09
C4-I2...O2	3.43	98.0	164.5	2.09
$[\text{I}_4\text{-1}^{2+}(\text{Cl}^-)(\text{TfO}^-)]$	$d_{\text{I}\cdots\text{Cl}}^a$	% vdW	$\angle \text{X-I}\cdots\text{Cl}^b$	$d_{\text{C}\cdots\text{I}}^a$
C3-I1...Cl1	3.08	82.6	174.7	2.10
C6-I2...Cl1	3.31	88.7	159.6	2.10
C17-I3...Cl1	3.13	83.9	167.7	2.10
C20-I4...Cl1	3.24	86.9	168.3	2.11
$[\text{I}_4\text{-1}^{2+}(\text{Cl}^-)_2]$	$d_{\text{I}\cdots\text{Cl}}^a$	% vdW	$\angle \text{X-I}\cdots\text{Cl}^b$	$d_{\text{C}\cdots\text{I}}^a$
C2-I1...Cl1	3.15	85.6	172.8	2.10
C4-I2...Cl1	3.31	89.9	160.9	2.11

^a In Å. ^b In $^\circ$.

to 17% shorter than the sum of the theoretical vdW radii (3.74 Å)¹⁷ (Fig. 2 and Table 1). This fact and the large angles ($\angle \text{C-I}\cdots\text{Cl} = 159.6^\circ\text{--}174.7^\circ$) clearly suggest directional halogen bonding. The tetra coordination of Cl^- certainly results in an overall strong XB complex thus displacing the probably weakly bonded triflates.

The asymmetric unit contains two different moieties of $\text{I}_4\text{-1}^{2+}$, a chloride anion, and a disordered triflate with short contacts to the aromatic pyridine protons. Infinite chains alternating between XB donors $\text{I}_4\text{-1}^{2+}$ and acceptors Cl^- are formed. The 1:1 stoichiometry for $\text{I}_4\text{-1}^{2+}$ and Cl^- confirms the key role of halogen bonding in this self-assembled structure. Triflate is here the second anion counterbalancing the overall charge of the dications. The pyridinium rings within each molecule are almost coplanar whereas the linking C=C double bond is strongly tilted of about 50° to 60° . In the structure of [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)_2$] the overall neutral dichloride complex appears to be assembled *via* weak $\text{CH}\cdots\text{Cl}^-$ hydrogen bonds. Details of the data collection and refinement are given in the ESI[†] (Tables S2-S10).

The strongly tilted conformation of the dications (tilt angles of about 50° to 60°) in the solid state is in agreement with the

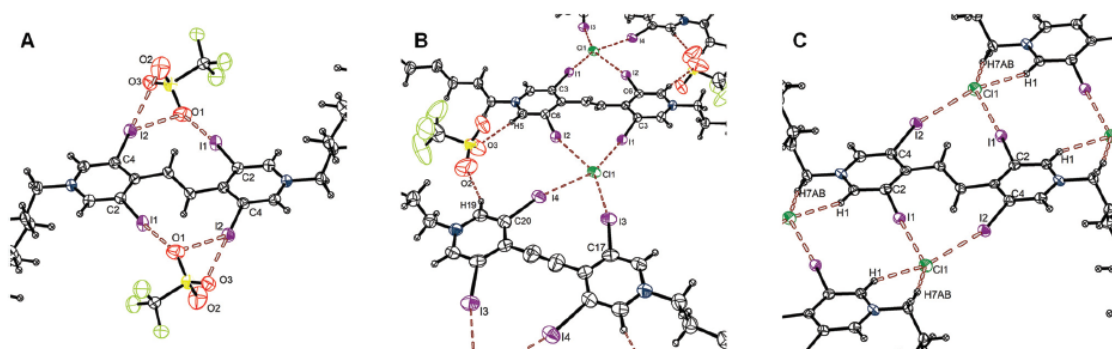


Fig. 2 Crystal structures of [$\text{I}_4\text{-1}^{2+}(\text{TfO}^-)_2$] (A), [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)(\text{TfO}^-)$] (B), and [$\text{I}_4\text{-1}^{2+}(\text{Cl}^-)_2$] (C). Ellipsoids are depicted at their 50% probability level. C – black, N – blue, O – red, F – light green, S – yellow, Cl – dark green, I – purple.

observation of a two electron transfer process during the electrochemical reduction of the dication. Even in homogeneous solution the steric hindrance of the relatively big halide substituents and the rigid H-C=C-H unit most probably lead to a conformation in which the heteroaromatic systems are not electronically coupled *via* the π -conjugated system and this even in the XB complex.

Additionally, the formation of the infinite alternating donor acceptor chains explains the solubility issues observed in ACN and DMSO. The relatively less polar ACN is incapable of separating the charges and upon chloride addition the polymeric complex precipitates. DMSO on the other hand is apparently able to form competitive XB to the viologens, thus occupying the second binding site, until getting replaced by the first chloride anion, after exceeding 1 eq.

2. DFT calculations

In order to obtain further information on the structures of I_4-1^{2+} and H_4-1^{2+} (as well as their onefold or twofold reduced analogues), DFT calculations (M06-2X D3 def2-TZVP, for further details and coordinates see the ESI[†]) were performed on model compounds in which all octyl groups were replaced by methyl ($I_4-1'^{2+}$ and $H_4-1'^{2+}$). Our main goal was to compare the geometries of these compounds and to check whether bidentate halogen bonding with halides was geometrically feasible for $I_4-1'^{2+}$, and thus the interpretation will focus entirely on those aspects (as we are well aware that the exact calculation of redox potentials and halogen bonding energies can be tricky). To further verify the performance of the applied theoretical model, we compared the bond lengths in $I_4-1'^{2+}$ with the ones of the cation in the crystal structure of $[I_4-1'^{2+}(\text{TfO}^-)_2]$, and deviations were (often far) below 0.01 Å. The optimized gas phase structures obtained for all six compounds ($H_4-1'^{2+}$, $H_4-1'^+$, H_4-1' as well as $I_4-1'^{2+}$, $I_4-1'^+$, I_4-1') clearly indicated that all bond lengths of the core structures were very similar (with deviations <0.1 Å) between two comparable species (*e.g.* $I_4-1'^{2+}$ and $H_4-1'^{2+}$, Fig. 3). However, there is a clear difference concerning the planarity when comparing the iodinated and the non-iodinated compounds. Due to the size of the four iodine substituents, all iodinated compounds deviate more or less strongly from planarity, while all non-iodinated compounds are more or less planar (with only $H_4-1'^{2+}$ featuring a slight tilt between the two planes of pyridines, tilt angle approx. 29° as measured by the I-C...C-I dihedral angle; see ESI[†]). In the iodinated compounds, $I_4-1'^{2+}$ shows the strongest deviation from planarity (tilt angle 72°), and similar to the progression

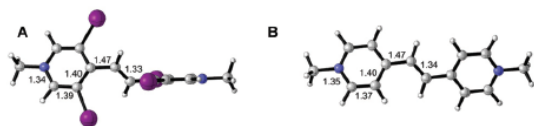


Fig. 3 Optimized gas phase structures of $I_4-1'^{2+}$ (A) and $H_4-1'^{2+}$ (B), with octyl groups replaced by methyl. Bond lengths in Å. Graphics by CYLview.¹⁶

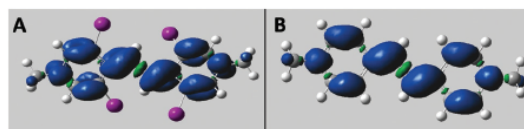


Fig. 4 Spin densities of $I_4-1'^{+}$ (A) and $H_4-1'^{+}$ (B).

in the series $H_4-1'^{m+}$ ($n = 2$ to 0), the structures of $I_4-1'^{+}$ (tilt angle 43°) and I_4-1' (tilt angle 17°) become more planar. In the latter case, this approach to planarity is accompanied by slight distortions of the carbon-iodine bonds, likely due to steric repulsion between the iodines (see ESI[†]). The spin densities of the radical intermediates $I_4-1'^{+}$ and $H_4-1'^{+}$ are depicted in Fig. 4. A coarse comparison of the two images indicates that the spin is similarly distributed in both compounds despite the differences in tilt angles (43° vs. 0°) mentioned above. The latter do not seem to have a strong influence on delocalization.

In the optimized structure of $I_4-1'^{2+}$ (Fig. 3 left), the electrophilic carbon-iodine axes on different pyridine moieties are almost perpendicular to each other, and thus a bidentate coordination of Lewis acids like halides seems improbable at first glance. In order to investigate the binding behavior further, the chloride, bromide and iodide complex of $I_4-1'^{2+}$ were optimized by the same means as mentioned above. All complexes feature a clear bidentate coordination of the halide, as is illustrated for chloride in Fig. 5, in which the two pyridine moieties of the halogen bond donor become almost co-planar.

There is, however, a slight distortion of the halogen bond donors, as the iodines which are coordinated to the halide are closer to each other than the ones on the opposite side of the molecule (3.58 Å vs. 4.06 Å for the chloride complex, Fig. 5). In all complexes, two slightly different halogen bonds are observed, with the stronger one always also being the slightly more linear one (see ESI[†]). Even though the determination of adduct energies was not our primary goal, a comparison of the iodine-halide distances seems to reveal a trend: compared to the respective sum of the van der Waals radii, the distance is shortest for the chloride complex (75% and 79% of van der Waals radii) and the longest for iodide (79% and 82%), with bromide in-between (77% and 80%). Thus, these gas-phase calculations seem to predict that chloride is most strongly bound by the halogen bond donors.

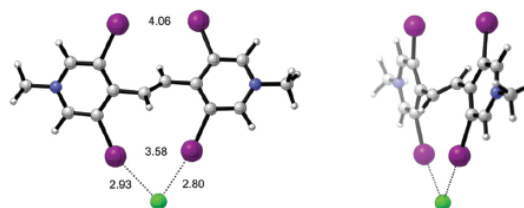


Fig. 5 Optimized gas phase structure of the complex of $I_4-1'^{2+}$ with chloride (front and side view). Distances for I-Cl and I-I in Å.

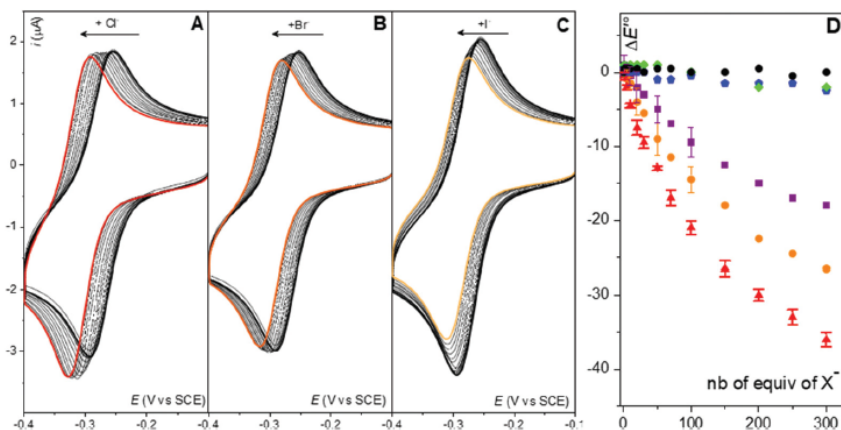


Fig. 6 CVs of 0.1 mM $\text{I}_4\text{-1}^{2+}$ in 0.1 M $n\text{Bu}_4\text{PF}_6/\text{DMF}$ in the absence and in the presence of increasing concentrations (from 0 to 30 mM) of $n\text{Bu}_4\text{X}$: (A) $\text{X} = \text{Cl}^-$, (B) $\text{X} = \text{Br}^-$ and (C) $\text{X} = \text{I}^-$. $T = 298\text{ K}$, $\nu = 0.1\text{ V s}^{-1}$. (D) Variation of the formal reduction potential of $\text{I}_4\text{-1}^{2+}$ as function of LB added (Cl^- : red, Br^- : orange, I^- : violet, NO_3^- : blue, HSO_4^- : green and OTf^- : black). The error bars indicated for chloride, bromide and iodide are standard deviations and each point is the average of at least three experiments.

3. XB detection

3.1 Cyclic voltammetry. The electrochemical behaviour of the three derivatives $\text{I}_4\text{-1}^{2+}$, $\text{Br}_4\text{-1}^{2+}$ and $\text{H}_4\text{-1}^{2+}$ was studied by CV in the presence of various anions. The DMF solutions contained 0.1 M $n\text{Bu}_4\text{PF}_6$ and in the titration experiments, the total salt concentration was kept constant never exceeding 0.11 M. In presence of 100 equivalents of chloride ($n\text{Bu}_4\text{Cl}$) the formal standard potential of the reduction wave of $\text{I}_4\text{-1}^{2+}$ was shifted toward more negative potentials ($\Delta E^{\circ'} = -21\text{ mV}$) (Fig. 6A).

Smaller but still significant potential shifts were recorded when adding bromide (-15 mV for 100 eq. Br^-) or iodide (-10 mV for 100 eq. I^-) (Fig. 6B and C and Table 2). These observations are consistent with a stabilization of the dicationic $\text{I}_4\text{-1}^{2+}$ species *via* charge-assisted non-covalent XB interactions leading to the complex $[\text{I}_4\text{-1}^{2+}, \text{Cl}^-]$. Such a reductive XB deactivation has been already demonstrated for iodo-phenylviologen derivatives in one of our previous articles.^{4,5} Reduction to the neutral state ($\text{I}_4\text{-1}$) proved to significantly decrease the XB-donor strength.

The shift intensity regarding the chosen halide can thus be directly correlated with the charge density of the XB-acceptors ($\text{Cl}^- > \text{Br}^- > \text{I}^-$). The less basic oxo-anions OTf^- , HSO_4^- or NO_3^- have been also tested with $\text{I}_4\text{-1}^{2+}$ but no significant perturbations of the CVs could be detected (Fig. S20, ESI[†]) suggesting a strong selectivity for halide anions. Replacing the iodo-substituents with bromine ($\text{Br}_4\text{-1}^{2+}$) or hydrogen ($\text{H}_4\text{-1}^{2+}$) resulted in receptors with no detectable halide affinity (Fig. S21 and S22, ESI[†]). These results indicate that under the experimental conditions $[\text{C-Br}\cdots\text{Cl}^-]$ XB interactions are too weak because of the lower polarizability of the XB-donor atom, decreasing for the homologous elements $\text{I} > \text{Br} > \text{Cl}$. Apparently $[\text{C-H}\cdots\text{Cl}^-]$ is also not competitive with the other non-covalent interactions in the polar liquid electrolyte.

The dependence of the observed formal potential shift $\Delta E^{\circ'}$ of $\text{I}_4\text{-1}^{2+}$ on the anion concentration is represented in Fig. 6D. The titration curves have been fitted according to eqn (1) based on a conventional square scheme mechanism (Scheme 5) assuming a 1:1 stoichiometry of the XB complexes according to the observed solid state crystal structure.

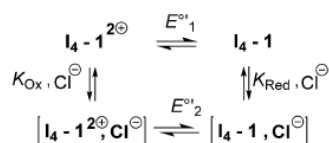
$$\Delta E^{\circ'} = E_2^{\circ'} - E_1^{\circ'} = -\frac{RT}{nF} \ln \frac{(1 + K_{\text{ox}}[\text{LB}])}{(1 + K_{\text{red}}[\text{LB}])} \quad (1)$$

With $E_1^{\circ'}$ being the formal standard potential of the redox couple $\text{I}_4\text{-1}^{2+}/\text{I}_4\text{-1}$ and $E_2^{\circ'}$ the nominal potential depending on the Lewis Base concentration $[\text{LB}]$ and K_{ox} and K_{red} the thermodynamic affinity constants for the interaction with $\text{I}_4\text{-1}^{2+}$ and $\text{I}_4\text{-1}$, respectively.

Table 2 Table of standard potential shifts ($\Delta E^{\circ'}$) for the redox couples dication/cation radical of $\text{I}_4\text{-1}^{2+}$, $\text{Br}_4\text{-1}^{2+}$ and $\text{H}_4\text{-1}^{2+}$ upon addition of $n\text{Bu}_4\text{X}$ ($\text{X} = \text{Cl}, \text{Br}, \text{I}, \text{OTf}, \text{HSO}_4$ and NO_3) and the corresponding affinity constants K_{ox} and K_{red}

$\text{X}_4\text{-1}^{2+}$	Anion	$\Delta E^{\circ'}/^a$	$\Delta E^{\circ'}/^b$	K_{ox}^c	K_{red}^c	$K_{\text{ox}}/K_{\text{red}}$
$\text{I}_4\text{-1}^{2+}$	Cl^-	-8	-21	255	60	4.2
	Br^-	-4	-15	135	35	3.8
	I^-	-2	-10	55	7	7.8
	OTf^-	0	0	—	—	—
	HSO_4^-	0	0	—	—	—
	NO_3^-	0	0	—	—	—
$\text{Br}_4\text{-1}^{2+}$	Cl^-	0	-4	14	6	2.3
	Br^-	0	-3	8	3	2.7
	I^-	0	0	—	—	—
$\text{H}_4\text{-1}^{2+}$	Cl^-	0	-5	—	—	—

^a In mV, upon addition of 20 equiv. of $n\text{Bu}_4\text{X}$, standard deviation: $\pm 1\text{ mV}$. ^b In mV, upon addition of 100 equiv. of $n\text{Bu}_4\text{X}$ standard deviation: $\pm 2\text{ mV}$. ^c In M^{-1} .



Scheme 5 Square scheme mechanism.

Fitting this square scheme model to the experimental titration plots allowed for the determination of the respective affinity constants depending on the oxidation state of the XB-donor. The highest affinity constant has been calculated for an association between $\text{I}_4\text{-1}^{2+}$ and Cl^- with a K_{ox} of 255 M^{-1} (Tables 2 and 3). As expected, the value of the calculated K_{red} is quite low, indicating only weak interactions between the neutral $\text{I}_4\text{-1}$ with chloride. The markedly weaker Lewis acidity of $\text{I}_4\text{-1}$ compared to $\text{I}_4\text{-1}^{2+}$ is also reflected by the size of their corresponding σ -holes, with a $V_{\text{s,max}}$ of 57 kJ mol^{-1} for $\text{I}_4\text{-1}$ and 584 kJ mol^{-1} for $\text{I}_4\text{-1}^{2+}$. When DFT computations including the SMD18¹⁹ solvation model for DMF were employed to estimate the binding constants, a similar trend was observed, with $K_{\text{ox}} = 62 \text{ M}^{-1}$ and $K_{\text{red}} = 1.6 \times 10^{-5} \text{ M}^{-1}$ (even though the difference between the two complexes was much more pronounced in this case, for currently unknown reasons).

The values of K_{ox} for bromide (135 M^{-1}) or iodide (55 M^{-1}) are consistent with the relative Lewis basicity. Low affinity constants could be determined for $\text{Br}_4\text{-1}^{2+}$ taking into account the small potential shifts. The addition of chloride did not show any effect neither on the first reduction wave of $\text{H}_4\text{-1}^{2+}$ nor on the second.

3.2 UV-vis spectroscopy. The results obtained with CV were compared with titration experiments employing UV-vis absorption spectroscopy. The spectrum of $\text{I}_4\text{-1}^{2+}$ was recorded in pure non-deuterated DMF at 298 K (Fig. 7A, black trace). Then the electrolyte salt $n\text{Bu}_4\text{PF}_6$ (0.1 M) was added in order to maintain comparable ionic strength to the electrochemical experiments. The maximum absorption peak centred at 334 nm was attributed to $\pi\text{-}\pi^*$ transitions.²⁰ As illustrated in Fig. 7A (red trace), upon addition of chloride ($n\text{Bu}_4\text{Cl}$), a decrease of the molar absorption coefficient for the band at 334 nm, accompanied with a slight bathochromic shift of 5 nm and the appearance of

Table 3 Association constants ($K_{\text{ox}}/\text{M}^{-1} \pm 6$)/binding free energies ($\Delta G/\text{kJ mol}^{-1} \pm 1$) of various anions with $\text{I}_4\text{-1}^{2+}$ determined by various techniques

Solvent	Anion	CV	UV-vis	¹ H NMR
0.1 M TBAPF ₆ in DMF	Cl ⁻	255 (-16.2)	187 (-15.3)	34 (-10.3)
	Br ⁻	135 (-14.4)	—	—
	I ⁻	55 (-11.7)	—	—
	TfO ⁻	nd	—	—
DMF	Cl ⁻	—	1466 (-21.4)	313 (-4.6)
	Br ⁻	—	365 (-17.3)	59 (-12.0)
	I ⁻	—	nd	66 (-12.3)
	TfO ⁻	—	nd	nd

nd: not determined.

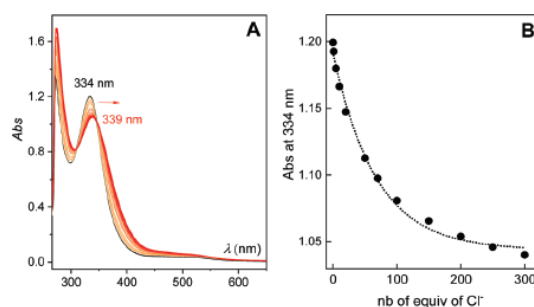


Fig. 7 (A) UV-vis spectra of 0.1 mM $\text{I}_4\text{-1}^{2+}$, in 0.1 M $n\text{Bu}_4\text{PF}_6/\text{DMF}$ in the presence of increasing equivalents of Cl^- (from 0, black trace up to 300 equivalents, red trace). (B) Variation of the absorbance at 334 nm as function of added equivalents of chloride. The dash curve was obtained by numerical simulation representing the best fit to the experimental data based on a 1:1 stoichiometry using *BindFit*.^{21,22}

an isobestic point. The association constant between $\text{I}_4\text{-1}^{2+}$ and Cl^- in the same solution as used for CV experiments was determined using *BindFit*, an online tool for supramolecular chemistry research and analysis.^{21,22} A binding constant K_a of 187 M^{-1} was calculated based on a 1:1 stoichiometry (Table 3). It is one order of magnitude smaller than the value obtained in the absence of $n\text{Bu}_4\text{PF}_6$ (1466 M^{-1} , Fig. S23, ESI[†]).

This observation highlights the significant competition of electrolyte ions at relatively high concentration (0.1 mM) with the species of interest in the following electrochemical experiments used to quantify halogen bonding.

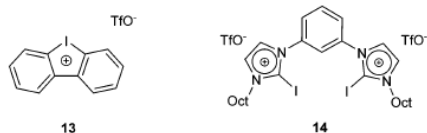
The association constant of $\text{I}_4\text{-1}^{2+}$ with Br^- in the absence of $n\text{Bu}_4\text{PF}_6$ was estimated from the UV-vis data (Fig. S24A, ESI[†]) to be $K_a = 365 \text{ M}^{-1}$ which is considerably smaller than the value obtained with Cl^- under the same conditions. The affinity constant for I^- could not be determined (Fig. S24B, ESI[†]), probably due to its very weak interaction with $\text{I}_4\text{-1}^{2+}$.

Titration with the weak Lewis base, triflate, did not result in any change of the absorption spectrum of $\text{I}_4\text{-1}^{2+}$ (Fig. S24D, ESI[†]).

Interestingly, the spectrum of the reference compound, the non-halogenated derivative $\text{H}_4\text{-1}^{2+}$, was not at all perturbed by the presence of the Lewis basic chloride at high concentrations (Fig. S25, ESI[†]), confirming the XB predominant interaction in the case of $\text{I}_4\text{-1}^{2+}$.

3.3 NMR spectroscopy. Taking advantage of the diamagnetic character of the dicationic receptors, ¹H NMR spectroscopy could also be used as a technique in order to determine affinity constants and to confront the results with those of the other techniques described above.

1 mM solutions of $\text{I}_4\text{-1}^{2+}$, $\text{Br}_4\text{-1}^{2+}$ and $\text{H}_4\text{-1}^{2+}$ in DMF and for comparison two previously described strong XB-donors, iodonium triflate 13²³ and bisimidazolium 14²⁴ (Scheme 6) were titrated with increasing amounts of chloride. ¹H NMR spectra of the chloride titration of $\text{I}_4\text{-1}^{2+}$ are shown in Fig. 8 (see also Fig. S26, ESI[†]). The signal of the pyridine protons of $\text{I}_4\text{-1}^{2+}$ is progressively shifting from 8.890 to 9.005 ppm upon the addition of about 83 equivalents of chloride. The corresponding



Scheme 6 Structures of the XB-donors iodolium **13**²³ and bisimidazolium **14**.²⁴

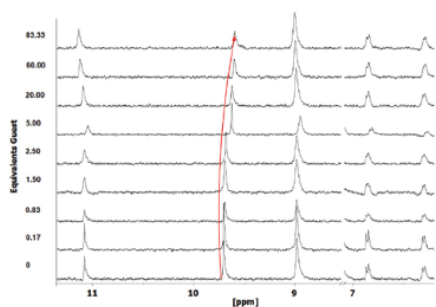


Fig. 8 NMR spectra of the chloride titrations of I_4-1^{2+} under standard suppression conditions. The red trace indicates the shift of the signals of interest. The signals at about 9 ppm and 11.1 ppm are residuals from the DMF. Those appear probably due to interactions with the XB-donors and are absent in pure DMF. The latter one starts shifting at higher ion concentrations.

association constants (1/1 stoichiometry) were determined by using the *BindFit* program.²²

With an association constant of 313 M^{-1} (*i.e.* $\Delta G = 14.6\text{ kJ mol}^{-1}$), I_4-1^{2+} exhibits the strongest binding to chloride of the herein studied bidentate derivatives (Fig. 9A). Br_4-1^{2+} displays a one order of magnitude smaller K_a value of about 28 M^{-1} . These results follow the classic trend of the increasing XB donor strength in the order $Br < I$.²⁵ For comparison, monodentate iodine(III)-based XB-donor **13** features a binding constant to chloride of 123 M^{-1} under the same conditions. Due to signal overlap, binding affinities could not be determined for

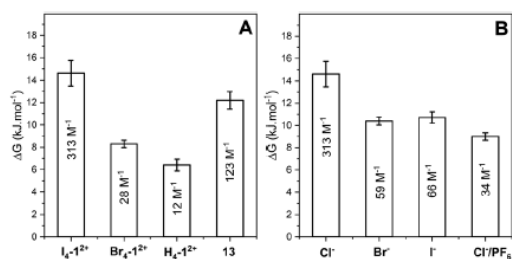


Fig. 9 Binding free energies ($kJ\ mol^{-1}$) and affinity constant (M^{-1}) determined by NMR-titration at 298 K in DMF-h7 for (1 mM) I_4-1^{2+} , Br_4-1^{2+} , H_4-1^{2+} , **13** towards chloride (left), and for I_4-1^{2+} towards chloride, bromide, iodide and towards chloride in the presence of 0.1 M nBu_4PF_6 (right). All compounds were fitted presuming a 1:1 stoichiometry. Association constants were calculated using the *BindFit* program.^{24–27}

bidentate XB donor **14** in this solvent. As expected the affinity constants for the interaction of I_4-1^{2+} with bromide (59 M^{-1}) and iodide (66 M^{-1}) are significantly smaller as with chloride (313 M^{-1}). The impact of the ionic strength of the solution was also controlled by comparing the recorded spectra during the titration with or without the supporting electrolyte nBu_4PF_6 .

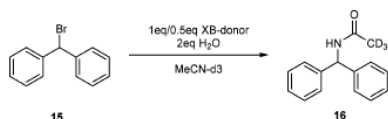
As expected and in accordance with electrochemical and UV-vis results, addition of supporting electrolyte induced a decrease of K_a by one order of magnitude (313 M^{-1} and 34 M^{-1} , respectively).

These results indicate the significant role of Coulombic interactions.

3.4 Discussion of the determined affinity constants. The three different techniques, CV, NMR and UV-vis spectroscopy, used for the determination of the binding free energies of X_4-1^{2+} towards halides are complementary. CV experiments require the presence of ionic species in form of a supporting electrolyte involving multiple competing electrostatic interactions in solution. UV-vis and NMR titrations can be realized in presence or in absence of such electrolyte salts, thus helping to discriminate the impact of XB among the different non-covalent interactions. Furthermore, electrochemistry allows determining elegantly the affinity constants of the reduced XB-donor species, named K_{red} in the discussed square scheme model (Scheme 5). These values are very difficult to obtain *via* UV-vis and NMR spectroscopy, requiring quantitative chemical reduction of the XB-donors prior to the titration experiments. Table 3 summarizes affinity constants and binding free energies for the interaction between I_4-1^{2+} and the four anions Cl^- , Br^- , I^- and OTf^- in DMF. It is noteworthy that the polar solvent DMF was chosen for solubility reasons, thus giving rise to relatively low formal affinity constants due to solvation of the interacting species. Dependency of the affinity constant on the nature of the solvent has been already reported²⁶ but its impact is difficult to explain and to predict.

All methods indicate the strongest interaction being between I_4-1^{2+} and chloride (Table 3), in agreement with the DFT calculations in the gas phase. The corresponding binding energy of Br_4-1^{2+} (-8.5 kJ mol^{-1}) is considerable weaker, which can be attributed to the weaker polarization of the Br substituents and is not experimentally measurable for H_4-1^{2+} . When comparing the data from the spectroscopic experiments (UV-vis and NMR) with $[I_4-1^{2+}, Cl^-]$ it can be stated that the competing electrostatic interactions of the supporting electrolyte salt significantly diminish the formal binding constants of up to one order of magnitude. In both systems, with and without nBu_4PF_6 , the order of binding strength was observed as follows: $Cl^- > Br^- \geq I^- \gg TfO^-$. These experimental results together with the obtained XRD structures for the complex $[I_4-1^{2+}, Cl^-]$ constitute further evidence of halogen bonding being the main driving force for the observed assemblies. Addition of triflate in the solution did not perturb the signals whatever the method, confirming the weak Lewis basicity of this anion.

When comparing the three employed analytical techniques the corresponding results obtained by CV and UV-vis certainly reflect the global electronic effects whereas NMR locally probes



Scheme 7 XB activated solvolysis of benzhydryl bromide to benzhydrylacetamide- d_3 .

distinct sites. However, all methods show the same tendency and confirm the highest affinity constant for the interaction between I_4-1^{2+} and Cl^- . To the best of our knowledge, such a systematic quantification of XB strength in solution using several analytical techniques has not been reported yet in literature.

4. Substrate activation/solvolyis

The determined binding constants are lower than expected for a dicationic chelating XB-donor, but are still high enough to consider potential catalytic activity in organic reactions. The catalytic properties of I_4-1^{2+} , Br_4-1^{2+} and H_4-1^{2+} were first tested in the Diels-Alder reaction between methylvinylketone and cyclopentadiene.²⁸ However, we do not observe any catalytic activity. As the activation of neutral substrates like carbonyl derivatives has proven to be more difficult than the activation of S_N1 -type substrates in halide abstraction reactions, we then turned our attention towards the latter. In this second approach, we investigated the bromide abstraction reaction of benzhydryl bromide in the presence of acetonitrile (Scheme 7).

As mentioned in the introduction, we had previously tested the 4,4'-azobis(halopyridinium) derivatives I_4-2^{2+} , Br_4-2^{2+} and H_4-2^{2+} in this benchmark reaction and we have demonstrated their ability to act as XB-donor catalysts.⁶ However, with these derivatives a side reaction occurred in which the halogen bond donors acted as oxidants towards bromide. One of the main goals of the present study was to prepare iodopyridinium-based bidentate activators which would be stable against halides.

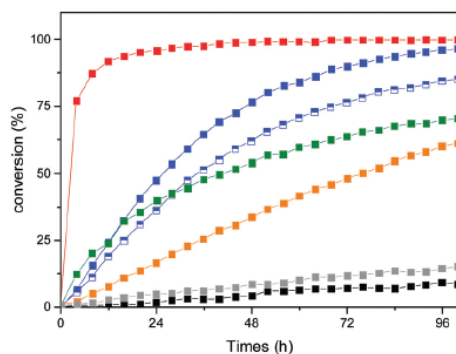


Fig. 10 Solvolysis of benzhydryl bromide with MeCN- d_3 and H_2O activated by the synthesized X_4-1^{2+} viologenes (I_4-1^{2+} in blue, Br_4-1^{2+} in orange, H_4-1^{2+} in grey and with 0.5 eq. of I_4-1^{2+} in blue/white) over 100 h at 22.5 °C. For comparison **13** is in red, **14** in green and the background in black.

The catalytic reaction was monitored *via* 1H NMR in MeCN- d_3 . A stoichiometric amount of I_4-1^{2+} reaches 97% conversion, while half of this amount leads to 86% (Fig. 10). In parallel, using Br_4-1^{2+} as XB-donor catalyst transforms 64% of the benzhydryl bromide and H_4-1^{2+} only 16%. The measured background over the same timespan conduces to 6% of transformation. The catalyst efficiency of I_4-1^{2+} and Br_4-1^{2+} is significantly higher than for the previously reported bis-imidazolium **4**.²⁹ The iodolium derivative **13**,³⁰ still displays the highest catalytic activity which is certainly due to the higher oxidation state of the iodine(III) center.

This result is in agreement with its high affinity constant to bromide in acetonitrile ($K_{a,30^\circ C} = 2.21 \times 10^6 M^{-1}$ (-8.8 kcal mol $^{-1}$)).^{30,31}

Conclusions

Bidentate halogen bond donors with halopyridinium moieties linked by an ethylene bridge were synthesized and structurally characterized by X-ray diffraction analysis and accompanying DFT calculations. Electrochemical measurements indicated clean twofold reductions to the corresponding neutral species, which should be much weaker XB-donors. Binding constants of the dicationic XB-donors to halides were obtained by three techniques: NMR, UV/vis and CV, with the latter also being able to estimate the binding strength of the reduced species – a clear advantage of this electrochemical approach. The data indicate that the presence of electrolyte salts in CV measurements leads to competition with the halides and a decrease of the observed binding constant by about an order of magnitude. Despite the strengths and weaknesses of the different methods, all three techniques revealed the same tendency of relative binding strength to halides, with chloride being the strongest. Compared to CV and UV-vis, NMR spectroscopy seems to significantly underestimate binding strength by almost one order of magnitude even when performed in the same experimental conditions (ionic strength, solvent, concentration...). We suggest that the results obtained with CV and UV-vis represent the effects on the global electronic structure of molecules, in contrast to NMR spectroscopy, which mainly probes local interactions. The Lewis acidity of the novel bidentate XB-donors was applied in a proof-of-principle case for halide abstraction, in which the chelating donors performed similarly compared to previously established bis(haloimidazolium) compounds. Furthermore, these new catalysts are chemically stable and did not decompose as their azo-bridged analogues.⁶ In contrast to our previously employed activators and catalysts, the new tetrahalo bipyridinium stilbenes are redox active, so that in principle it should be possible to control their catalytic activity by reversible switching the redox state. Future work on this interesting concept is in progress.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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Notes and references

- 1 S. Groni, T. Maby-Raud, C. Fave, M. Branca and B. Schöllhorn, *Chem. Commun.*, 2014, 50, 14616.
- 2 (a) R. Oliveira, S. Groni, C. Fave, M. Branca, F. Mavre, D. Lorcy, M. Fourmigué and B. Schöllhorn, *Phys. Chem. Chem. Phys.*, 2016, 18, 15867; (b) R. Oliveira, S. Groni, A. Vacher, F. Barrière, D. Lorcy, M. Fourmigué, E. Maisonhaute, B. Schöllhorn and C. Fave, *ChemistrySelect*, 2018, 3, 8874.
- 3 (a) J. Y. C. Lim, M. J. Cunningham, J. J. Davis and P. D. Beer, *Chem. Commun.*, 2015, 51, 14640; (b) J. Y. C. Lim and P. D. Beer, *Eur. J. Inorg. Chem.*, 2017, 220.
- 4 G. Creste, S. Groni, C. Fave, M. Branca and B. Schöllhorn, *Faraday Discuss.*, 2017, 203, 301.
- 5 C. Fave and B. Schöllhorn, *Curr. Opin. Electrochem.*, 2019, 15, 89.
- 6 F. Kniep, S. M. Walter, E. Herdtweck and S. M. Huber, *Chem. – Eur. J.*, 2012, 18, 1306.
- 7 A. M. Allgeier and C. A. Mirkin, *Angew. Chem., Int. Ed.*, 1998, 37, 894.
- 8 J. Choudhury, *Tetrahedron Lett.*, 2018, 59, 487.
- 9 (a) T. M. Beale, M. G. Chudzinski, M. G. Sarwar and M. S. Taylor, *Chem. Soc. Rev.*, 2013, 42, 1667; (b) C. Laurence, J. Graton, M. Berthelot and M. J. El Ghomari, *Chem. – Eur. J.*, 2011, 17, 10431; (c) A. Caballero, F. Zapata, N. G. White, P. J. Costa, V. Félix and P. D. Beer, *Angew. Chem., Int. Ed.*, 2012, 51, 1876.
- 10 (a) E. Gleich and Z. Warnket, *Phosphorus, Sulfur Silicon Relat. Elem.*, 1991, 55, 9; (b) K. A. Chudov, K. S. Levchenko, N. O. Poroshin, A. V. Shchegolkov, P. S. Shmelin and E. P. Grebennikov, *Russ. Chem. Bull., Int. Ed.*, 2019, 68, 1565.
- 11 The chlorinated analogs were synthesized for the sake of completeness, but due to the low halogen bonding capabilities of the bromo derivatives never investigated any further.
- 12 A. Krasovskiy, V. Krasovskaya and P. Knochel, *Angew. Chem., Int. Ed.*, 2006, 45, 2958.
- 13 J. Gorecka, C. Heiss, R. Scopelliti and M. Schlosser, *Org. Lett.*, 2004, 6, 4591.
- 14 J. J. Baldwin, K. Mensler and G. S. Ponticello, *J. Org. Chem.*, 1978, 43, 4878.
- 15 I. Sasaki, J. C. Daran and G. G. A. Balavoine, *Synthesis*, 1999, 815.
- 16 S. Amaboldi, A. Gennaro, A. A. Isse and P. R. Mussini, *Electrochim. Acta*, 2015, 158, 427.
- 17 (a) A. Bondi, *J. Phys. Chem.*, 1964, 68, 441; (b) S. C. Nyburg and C. H. Faerman, *Acta Crystallogr., Sect. B: Struct. Sci.*, 1985, 41, 274.
- 18 C. Y. Legault, *CYLview, 1.0b*, Université de Sherbrooke, 2009, <http://www.cylview.org>.
- 19 E. Engelage, N. Schulz, F. Heinen, S. M. Huber, D. G. Truhlar and C. J. Cramer, *Chem. – Eur. J.*, 2018, 24, 15983.
- 20 J. Deng, X. Fu, G. Wang, L. Wu and J. Hiang, *Electrochim. Acta*, 2012, 85, 195.
- 21 P. Thordarson, *Chem. Soc. Rev.*, 2011, 40, 1305.
- 22 <http://app.supramolecular.org/BindFit/>.
- 23 F. Heinen, E. Engelage, A. Dreger, R. Weiss and S. M. Huber, *Angew. Chem., Int. Ed.*, 2018, 57, 3830.
- 24 F. Kniep, L. Rout, S. M. Walter, H. K. V. Bensch, S. H. Jungbauer, E. Herdtweck and S. M. Huber, *Chem. Commun.*, 2012, 48, 9299.
- 25 (a) P. Metrangolo, F. Meyer, T. Pilati, G. Resnati and G. Terraneo, *Angew. Chem., Int. Ed.*, 2008, 47, 6114; (b) A. C. Legon, *Phys. Chem. Chem. Phys.*, 2010, 12, 7736; (c) A. Priimagi, G. Cavallo, P. Metrangolo and G. Resnati, *Acc. Chem. Res.*, 2013, 46, 2686.
- 26 S. M. Walter, F. Kniep, L. Rout, F. P. Schmidtchen, E. Herdtweck and S. M. Huber, *J. Am. Chem. Soc.*, 2012, 134, 8507.
- 27 D. Brynn Hibbert and P. Thordarson, *Chem. Commun.*, 2016, 52, 12792.
- 28 S. H. Jungbauer, S. M. Walter, S. Schindler, L. Rout, F. Kniep and S. M. Huber, *Chem. Commun.*, 2014, 50, 6281.
- 29 S. M. Walter, F. Kniep, E. Herdtweck and S. M. Huber, *Angew. Chem., Int. Ed.*, 2011, 50, 7187.
- 30 F. Heinen, E. Engelage, A. Dreger, R. Weiss and S. M. Huber, *Angew. Chem., Int. Ed.*, 2018, 57, 3830.
- 31 E. Engelage, N. Schulz, F. Heinen, S. M. Huber, D. G. Truhlar and C. J. Cramer, *Chem. – Eur. J.*, 2018, 24, 15983.