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# Electronic-Specific Modeling of Nitric Oxide in a Recombining Air Plasma

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**An electronic-specific collisional-radiative (CR) model of Nitric Oxide (NO) is presented and applied to the study of a recombining air plasma. This plasma is produced at local thermodynamic equilibrium at 8000 K and 1 atm by a 50-kW Inductively Coupled Plasma torch and passes at high-velocity through a water-cooled tube that forces rapid cooling and recombination. The electronic-specific model takes into account NO( $C^2\Pi$ ) predissociation, spontaneous emission from the  $C^2\Pi$  and  $D^2\Sigma^+$  states towards the  $A^2\Sigma^+$  state, and quenching of the excited states by various colliders. This model, alongside two other electronic-specific CR models developed for NO at NASA, is compared to measurements of NO excited states and electron densities along the tube length. Better agreement is obtained with the model presented in this work, especially because of the NO( $C^2\Pi$ ) predissociation and NO  $C^2\Pi$  and  $D^2\Sigma^+$  spontaneous emission toward NO( $A^2\Sigma^+$ ). The electron overpopulation is also explained by the partial equilibrium between N and O atoms and electrons, and the slow depletion of N and O atoms by the NO recombination reactions.**

## I. Introduction

**D**uring atmospheric entry, a vehicle experiences a significant radiative heat flux due to the shock layer that forms in front of the capsule. Designing the heat shield requires to accurately determine the heat load on the spacecraft. While the prediction of the radiative heat flux to the forebody has been thoroughly investigated, the prediction of radiation on the afterbody still suffers from large uncertainties [1,2]. Indeed, the recombination kinetics in the afterbody region are complex and lead to non-equilibrium population distributions of the strongest radiative species. This work will focus on these kinetics: the slow recombination reaction cannot accommodate the rapid hydrodynamic cooling of the gas in the afterbody wake, which leads to a departure from thermochemical equilibrium.

Experiments have been performed to acquire data representative of such recombining flows [3–6]: a plasma torch was used to heat up various mixtures of gas up to 7000-8000 K at equilibrium, then the gas was forced to cool rapidly while flowing through a water-cooled tube and various diagnostics were applied to quantify the non-equilibrium state of the gas at the outlet of tubes of various lengths, ranging from 10 to 65 cm. These measurements are valuable to assess the state-of-the-art kinetic models. In [7,8], two molecular and atomic nitrogen state-to-state models were assessed against the experimental data with a nitrogen/argon mixture. The present paper extends the assessment to an air/argon mixture.

The air-argon mixture experiments were performed at Stanford and are described in Ref. [4]: at the inlet of the tube (outlet of the plasma torch) the air/argon mixture is at LTE at 8000 K and atmospheric pressure, then the water-cooled tube forces plasma cooling to 2500 K. The decrease in temperature is sufficiently fast (from 8000 to 2500 K in approximately 2 ms) that the gas is out of equilibrium at the outlet of the tube. The NO excited electronic states

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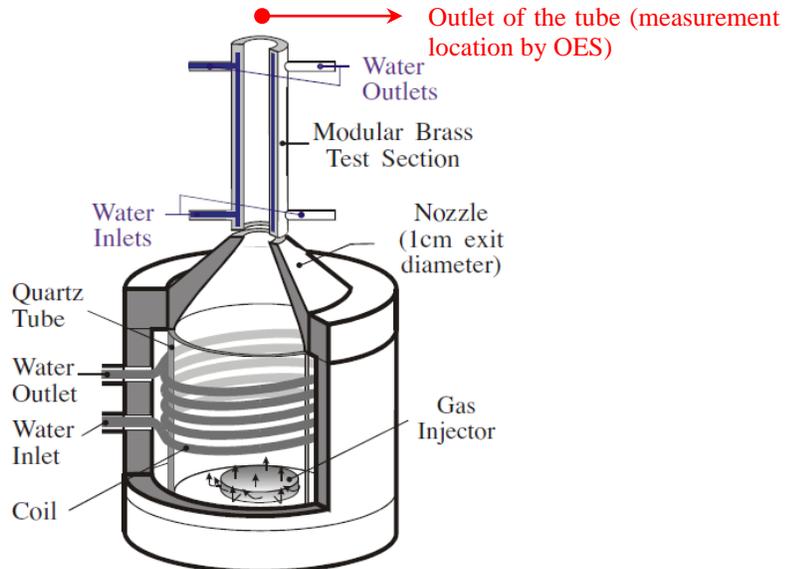
were found to be strong emitters of the recombining plasma and their number densities were measured by OES [4]. These measurements will serve as a reference in this work.

The widely used Park kinetic model [9] was compared to the experimental results in [4]. This model tracks the evolution of the global species and the internal states are assumed to follow Boltzmann distributions. Meanwhile, new electronic-specific kinetic models for air [10,11] were developed and validated against NASA Electric Arc Shock-Tube (EAST) measurements in a dissociating flow. But they have not yet been assessed in a recombining flow.

In this study, we will assess the NO electronic-specific kinetic model of Refs. [10,11] by comparing their predictions to the experimental overpopulations of NO electronic states measured in the air-argon recombination experiment [4]. Then, we will present a new electronic-specific CR model for Nitric Oxide that is in better agreement with the measurements.

## II. Experimental setup

The experimental setup is described in [4,12]. A dissociated high-temperature plasma is forced to recombine by rapid cooling. The equilibrium air/argon plasma is generated using a 50-kW ICP plasma torch facility. Then the plasma flows through a water-cooled tube (diameter 1cm, length ranging from 10 to 65 cm) to impose fast cooling (see Figure 1).



**Figure 1 Schematic of the water-cooled tube (at the top) and plasma torch.**

The mass flow rates at the entrance of the torch are 15.6 slpm of air, 162 slpm of argon and 2.3 slpm of H<sub>2</sub>. At the inlet of the tube (outlet of the torch) the gas is in thermal and chemical equilibrium at approximately 8000 K. It leaves the tube at approximately 2500 K. Emission spectroscopy was performed at the tube outlet (see Figure 1) for different tube lengths: 0, 10, 15, 40, 50 and 65 cm. The temperatures were measured using OH A-X emission bands around 308 nm, and the electron densities were obtained using the Stark broadening of the H<sub>β</sub> line, as described in Ref. [4]. The measured and equilibrium electron densities are given in Table 1. Overpopulation factors, defined in Eq. (1) as the ratio between the density of the species M and its equilibrium density, are also given in this table.

$$\rho_M = \frac{[M]}{[M]_{eq}} \quad (1)$$

**Table 1 Electron densities (measured and at equilibrium) and measured electron overpopulation factor at the exit of tubes of different lengths z [4].**

z [cm]	0	10	15	40	50	65
T [K]	7980 ± 144	6687 ± 144	5863 ± 125	3450 ± 150	2800 ± 100	2510 ± 100

$n_e^{eq}$ [cm <sup>-3</sup> ]	$1.5 \pm 0.3 \cdot 10^{15}$	$2.7 \pm 0.8 \cdot 10^{14}$	$4.6 \pm 1.0 \cdot 10^{13}$	(7.1 – 40) · 10 <sup>10</sup>	$8.4 \cdot 10^8 - 1.7 \cdot 10^{10}$	$4.5 \cdot 10^7 - 2.0 \cdot 10^9$
$n_e^{exp}$ [cm <sup>-3</sup> ]	$1.5 \pm 0.03 \cdot 10^{15}$	$2.6 \pm 0.1 \cdot 10^{14}$	$1.4 \pm 0.06 \cdot 10^{14}$	(3.4 – 20) · 10 <sup>11</sup>	$5.9 \cdot 10^{10} - 1.2 \cdot 10^{12}$	$2.3 \cdot 10^{10} - 1.1 \cdot 10^{12}$
$\rho_e$ [-]	0.8 – 1.3	0.7 – 1.4	2.4 – 4.1	0.8 – 28	$3.5 - 1.4 \cdot 10^3$	$12 - 2.4 \cdot 10^4$

The spectra also provided insights into the overpopulation factors of NO electronic states A<sup>2</sup>Σ<sup>+</sup>, B<sup>2</sup>Π, C<sup>2</sup>Π, and D<sup>2</sup>Σ<sup>+</sup>. The NO(C<sup>2</sup>Π) and NO(D<sup>2</sup>Σ<sup>+</sup>) are strongly overpopulated at 40, 50 and 65 cm owing to chemical nonequilibrium. Table 2 contains the overpopulation factors of NO(A<sup>2</sup>Σ<sup>+</sup>) and NO(C<sup>2</sup>Π) obtained from the measured spectra at various lengths.

**Table 2 NO(A,B,C,D) overpopulation factor at the exit of tubes of different lengths.**

z [cm]	0	10	15	40	50	65
$\rho_{NO(A,B)}$ [-]	1	1	1	8	500	$3.8 \cdot 10^4$
$\rho_{NO(C,D)}$ [-]	1	1	1	25	4600	$2.7 \cdot 10^5$

The overpopulation factors of the N<sub>2</sub> C<sup>3</sup>Π<sub>u</sub> state have also been measured by Gessman *et al.* [4,12] and are given in Table 3 for different tube lengths. The departure from equilibrium is very significant for tube lengths longer than 40 cm, with overpopulations of N<sub>2</sub>(C) between 10<sup>5</sup> and 10<sup>11</sup>.

**Table 3 N<sub>2</sub>(C) overpopulation factor at different lengths of the tube.**

z [cm]	0	10	15	40	50	65
$\rho_{N_2(C)}$ [-]	1	1	1	$2.5 \cdot 10^5$	$7 \cdot 10^8$	$3.5 \cdot 10^{11}$

### III. Kinetic model and simulation parameters

The kinetic model used is based on Park’s mechanism [9]. We modified the rates of certain reactions, using the modified rate coefficients proposed by Gessman [12] and those of Laux *et al.* [13] for reactions with Argon. We do not use the corrected coefficient of Gessman for the dissociation of N<sub>2</sub> by N ( $N_2 + N \rightarrow N + N + N$ ), since it was not experimentally deduced. We also used coefficients taken from Gorelov *et al.* [14]. The complete model (with the electronic-specific reactions) is detailed in the Appendix. It considers 19 species (and pseudo-species): Ar, Ar<sup>+</sup>, N, N<sup>+</sup>, N<sub>2</sub>(X,A,B), N<sub>2</sub><sup>+</sup>, O, O<sup>+</sup>, O<sub>2</sub>, O<sub>2</sub><sup>+</sup>, NO(X,A,B,C,D), NO<sup>+</sup> and electrons. This model only uses one temperature  $T$  because the high pressure (atmospheric pressure) prevents any departure from thermal equilibrium (the electron translational temperature is equal to the gas temperature) [15].

The complete list of the model’s reactions with their parameters is given in Table 5 of the Appendix.

#### A. NO electronic states modeling

##### 1. NO(C) predissociation

The NO C<sup>2</sup>Π state predissociates into N and O atoms, and NO can also be produced in the C<sup>2</sup>Π state through inverse predissociation. This phenomenon has been studied in Ref. [16]. We use the same predissociation rate as in Ref. [16],  $A_{NO(C)}^{pred} = 3.3 \cdot 10^8 \text{ s}^{-1}$ .

##### 2. NO electronic states radiative lifetime

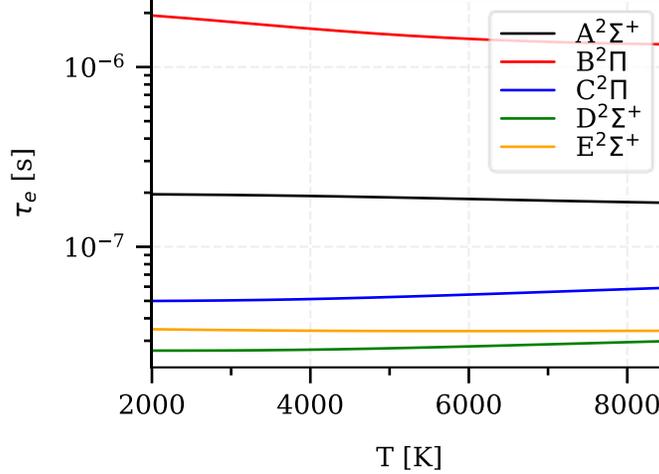
The radiative lifetime of an electronic level  $e$  is calculated in Eq. (2):

$$\frac{1}{\tau_e} = \sum_{v'} \sum_{v''} \frac{A_{v'v''} \exp\left(-\frac{E_{v'}}{kT}\right)}{Q_{vib}(T)} \quad (2)$$

Using the Einstein coefficients  $A_{v'v''}$  computed by Laux [16], we obtained the electronic radiative lifetimes of NO electronic levels, as displayed on Figure 2. These lifetimes correspond to the transitions from electronic states A,B,C,D,E to the ground state X only: for the C→A transition, we used Groth *et al.* rate constant [17], which leads to a radiative lifetime of  $\tau_{C \rightarrow A} = 7.4 \cdot 10^{-8} \text{ s}$ . For the D → A transition, we assume that since the C and D levels are very

close, the electron transition moment functions (ETMF) of the  $C \rightarrow A$  and  $D \rightarrow A$  transitions are similar and that the only difference will come from the proportionality  $\alpha$  factor (given in Eq. (3)) which is 2 for D and 1 for C, leading to  $\tau_{C \rightarrow A} = 2\tau_{D \rightarrow A}$  (cf. page 106 of Ref. [16]).

$$\alpha = \frac{2 - \delta_{0,\Lambda'+\Lambda''}}{2 - \delta_{0,\Lambda'}} \quad (3)$$



**Figure 2 Radiative lifetimes of NO electronic levels as function of the temperature.**

The radiative lifetimes were approximated constant over the temperature range. We obtained important differences with Johnston’s radiative lifetimes [10], since he used escape factors to model the autoabsorption in the shock-tube. In this work, as a first step, we considered an optically-thin medium and thus escape factors equal to 1.

### 3. Quenching of NO electronic states

The populations of NO electronic states are strongly affected by collisional quenching. However, while most models consider collisional excitation or quenching as independent of the collider [10,14], the quenching rate coefficients depend on the collider. Raiche and Crosley [18] have seen a dependence on the collision frequency (*i.e.* on  $\sqrt{T}$ ) for the quenching rate coefficients of NO(A) and NO(B) between 300 and 750 K. This approximation is also used by Park in Ref. [19], where he reviews several experimental quenching data (which are also considered below). If experimental data are not available, we will use Eq. (4), even for temperatures up to 8000 K:

$$k_q = A \cdot \sqrt{T} \quad (4)$$

Quenching rate coefficients of NO(A) and NO(D) are given in Table 4 for different colliders (Ar, NO(X), N<sub>2</sub> and O<sub>2</sub>). The adjusted  $A$  coefficients are obtained by fitting experimental data from 300 to 4500 K for NO(A), and at 300 K for NO(D) (no high-temperature measurements for NO(D)). For NO(A), the adjusted coefficients values are given below. For NO(D), we used experimental data from [20–28]. The quenching of NO(D) to NO(C) by collisions with atoms has also been studied by Laux (Eq. (5.48), page 108 of [16]) and uses the formula given in Eq. (5):

$$k_{coll}^{D \rightarrow C} [M] [s^{-1}] = 0.6 \cdot 10^{10} \sqrt{\frac{300}{T}} P[atm] X_M \quad (5)$$

with  $X_M$  the mole fraction of the species M in the mixture.

**Table 4 Adjusted coefficients for quenching of NO(A) and NO(D).**

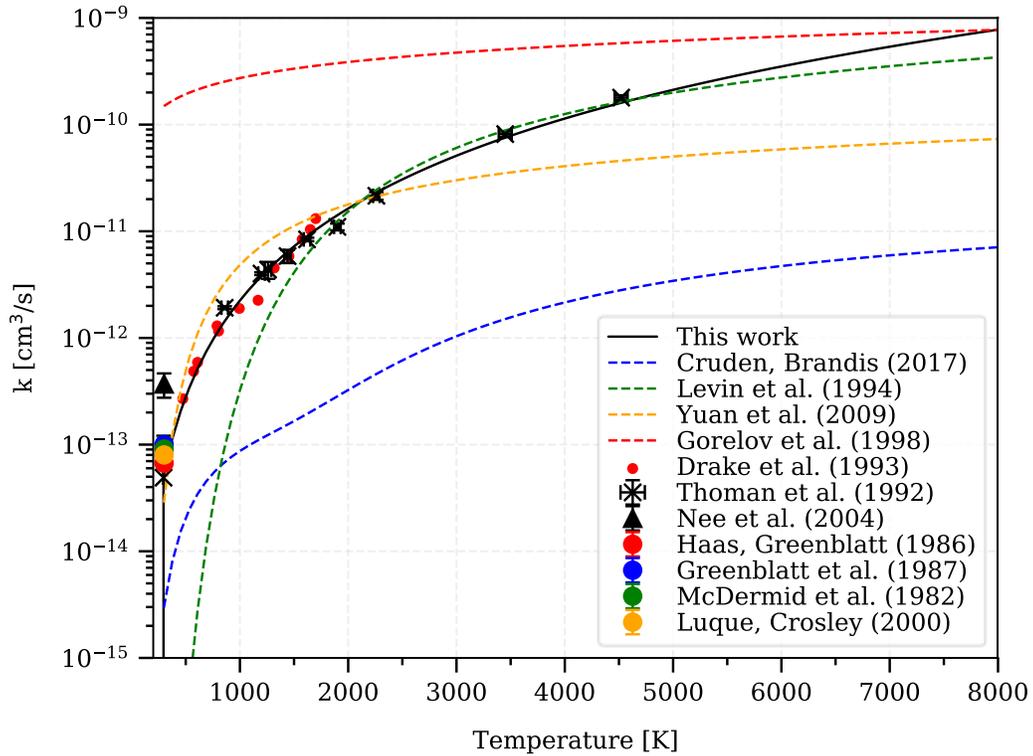
Quencher	$A$ [cm <sup>3</sup> /s]	
	NO(A)	NO(D)
M (Gorelov <i>et al.</i> [14])	$8.63 \cdot 10^{-12}$	$1.26 \cdot 10^{-11}$
Ar	$5.75 \cdot 10^{-15}$	$8.72 \cdot 10^{-12}$
NO(X)	$1.44 \cdot 10^{-11}$	$2.09 \cdot 10^{-11}$

N <sub>2</sub>	—	8.63 · 10 <sup>-12</sup>
O <sub>2</sub>	8.63 · 10 <sup>-12</sup>	3.49 · 10 <sup>-11</sup>

The quenching of NO(A) by N<sub>2</sub> can be fitted using an Arrhenius law over Thoman *et al.* data [29], using the rate coefficient given in Eq. (6):

$$k_q^{N_2} = 1.67 \cdot 10^{-20} \cdot T^{2.74} \cdot \exp\left(-\frac{182.8}{T}\right) \text{ cm}^3/\text{s} \quad (6)$$

This fit matches well the experimental data of [22,30–34] over a large temperature range (300 K – 8000 K). The other fits suggested by [11,35,36] do not agree as well over the large range of temperatures considered. The collisional quenching rate constant of NO(A) by N<sub>2</sub> is shown in Figure 3.



**Figure 3 Rate coefficients of NO(A) quenching by N<sub>2</sub>.**

The other quenching rate coefficients are displayed on Figure 4. Regarding the quenching by Argon, experimental measurements at low temperature (lower than 1500 K) are scattered over a few orders of magnitudes [30,31,33,34,37,38]. We adjusted the *A* coefficient to Gray *et al.* [38] data at higher temperatures (up to 2000 K approximately). This reaction may be important for the population of the NO A<sup>2</sup>Σ<sup>+</sup> state: indeed, the mixture is made of approximately 90% of Argon, which thus will be the main collider. Even though the quenching efficiency of Argon is small compared to other species (e.g. NO or O<sub>2</sub>), the preponderance of Argon in the mixture may make this reaction one of the most important regarding NO(A) depletion. The rate coefficients given labelled as from Gorelov *et al.* [14] and as from Cruden and Brandis [11] are taken from their coefficients of quenching by an undisciminated collider M. It is assumed that this rate used in [11,14] for species present in air also applies to Argon (and to Hydrogen). The simulations performed using Cruden and Brandis [11] and Johnston’s [10] models will consider Argon as a regular collider.

The quenching rate coefficients of NO(A) by H<sub>2</sub> are shown here in prevision of future works including the trace of Hydrogen present in the mixture. Since H<sub>2</sub> densities are very low compared to the other species, it can be safely neglected. However, the quenching rate coefficient varies of almost 3 orders of magnitude between 300 and 2000 K.

The rate coefficients given by [11,14,35] for the quenching by NO are close to each other and in correct agreement with the experimental data from [22,25,30,37,38]. We modified the  $A$  coefficient from Cruden and Brandis [11] to fit more closely the high-temperature measurements.

Regarding the quenching by  $O_2$ , the coefficients given by [11,14,35,36] are similar and agree well with the experimental data from [22,23,30–34,37,39] for temperatures up to 2500 K. This model uses the rate coefficient of Gorelov *et al.* [14] since it agrees well with the experimental data at the highest temperatures.

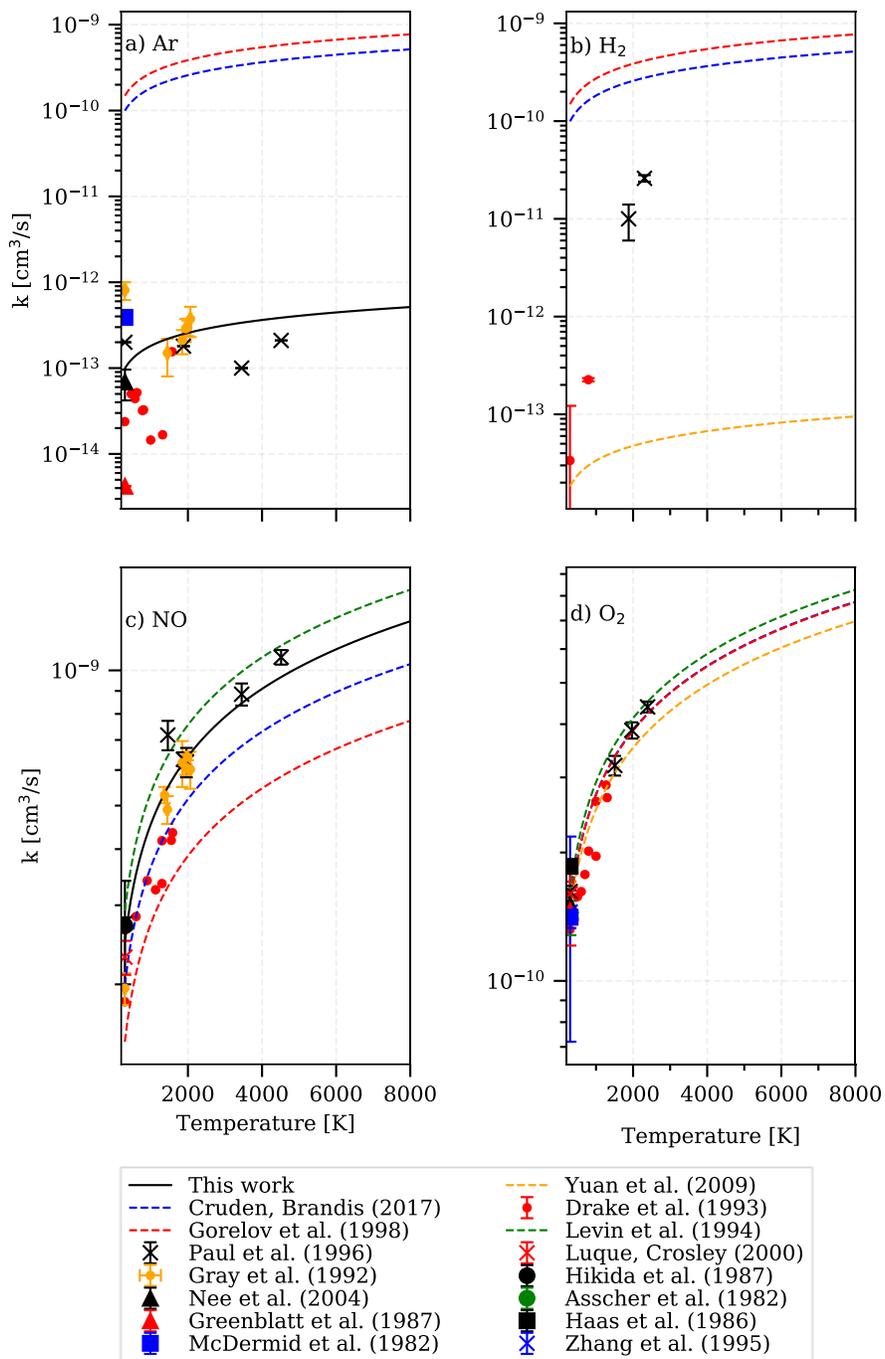


Figure 4 Rate coefficients of NO(A) quenching by (a) Argon, (b) H<sub>2</sub>, (c) NO and (d) O<sub>2</sub>.

Regarding the quenching of NO(B), a few experimental data at low-temperatures (300 K) are available [18,40–42] for various colliders (Argon, H<sub>2</sub>, N<sub>2</sub>, O<sub>2</sub> and NO). These points are scattered over orders of magnitude for the same collider. We chose to use the same rate coefficients as [11], which are taken from [43] and are in good agreement with Raiche *et al.* [18] for the quenching by O<sub>2</sub> up to 1000 K.

Imajo *et al.* [24] measured quenching rate coefficients for NO(C) by N<sub>2</sub> and NO at 300 K: we use  $A^{N_2} = 3.3 \cdot 10^{-10} \text{ cm}^3/\text{s}$  and  $A^{NO} = 13 \cdot 10^{-10} \text{ cm}^3/\text{s}$ . Although the branching ratio between C and D could explain the difference with the quenching coefficients by NO, it does not explain the differences with the quenching rate coefficients by N<sub>2</sub>. There might not be a correlation between C and D quenching rates, even though these two electronic states are close to each other. For the quenching of NO(C) by Argon and O<sub>2</sub>, we used the rate coefficients used by [11] and taken from [43].

For the quenching of the A, B, C and D states by atoms, we considered a fraction (0.6) of the kinetic rate, as was done for the quenching of D to A:  $k_q^{atom} = 0.6 \cdot 10^{-10} \cdot \sqrt{T} \text{ cm}^3/\text{s}$ . In future work, we will study in more details the sensitivity of the results to these quenching rate coefficients by atomic species.

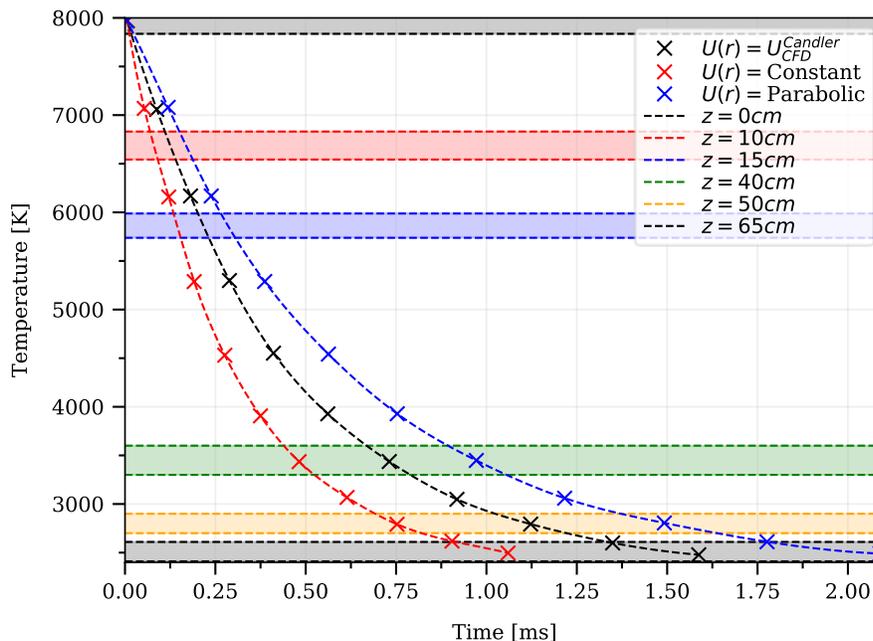
#### 4. $N_2(A) + NO(X) \rightarrow N_2(X) + NO(A)$

The quenching of N<sub>2</sub>(A) by NO can excite NO into the A<sup>2</sup>Σ<sup>+</sup> state. This reaction is important for the population of NO(A). Experimental data have been obtained at low temperatures (196 or 300 K) for N<sub>2</sub>(A) quenching by NO [44]. We used the rate coefficients given by Levin *et al.* [35] for this reaction, considered constant over the temperature range.

However, it requires to know the population of N<sub>2</sub> A<sup>3</sup>Σ<sub>u</sub><sup>+</sup> electronic state. We used the reaction rate constants given by Gorelov *et al.* [14] for N<sub>2</sub> electronic states X, A and B.

## B. Simulation parameters

To simulate the water-cooled tube, we performed 0D simulations at constant pressure (1 atm), using the temperature evolution along the centerline of the tube. Figure 5 shows the temperature evolutions used in our simulations. They were obtained using 3 different assumptions on the radial temperature profile in the tube: a constant radial temperature, a parabolic radial temperature, or a radial profile obtained with CFD simulation by Candler *et al.* [15]. We do not consider self-absorption and do not take into account species diffusion. To compute the equilibrium populations, we used the NASA CEA code [45] at the corresponding pressure (1 atm) and temperatures. Then we can use these equilibrium populations to calculate the overpopulation factors.



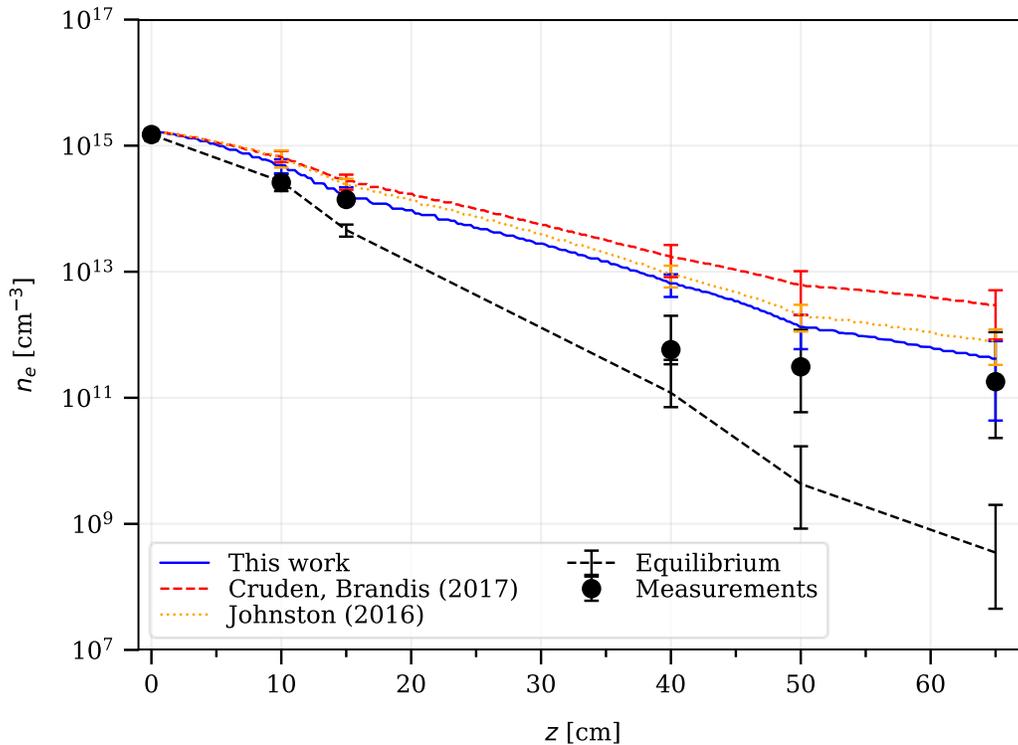
**Figure 5** Temperature evolutions used for the kinetic simulations.

The models of Johnston [10] and of Cruden and Brandis [11] have been slightly modified to take into account the high density of argon in the mixture: we added the ionization rate coefficients of Argon by electron impact from Laux *et al.* [13], as well as the charge exchange reaction between Argon and other ions. The rate constants of the charge exchange reaction  $\text{Ar}^+ + \text{M} \rightarrow \text{Ar} + \text{M}^+$  are taken equal to the rate coefficient of the charge exchange reaction between Argon and  $\text{N}_2$  given in Laux *et al.* [13]. Moreover, when collisions with third bodies were given in both models, we took the same third-body efficiencies for Argon as for other heavy particles.

#### IV. Results and discussion

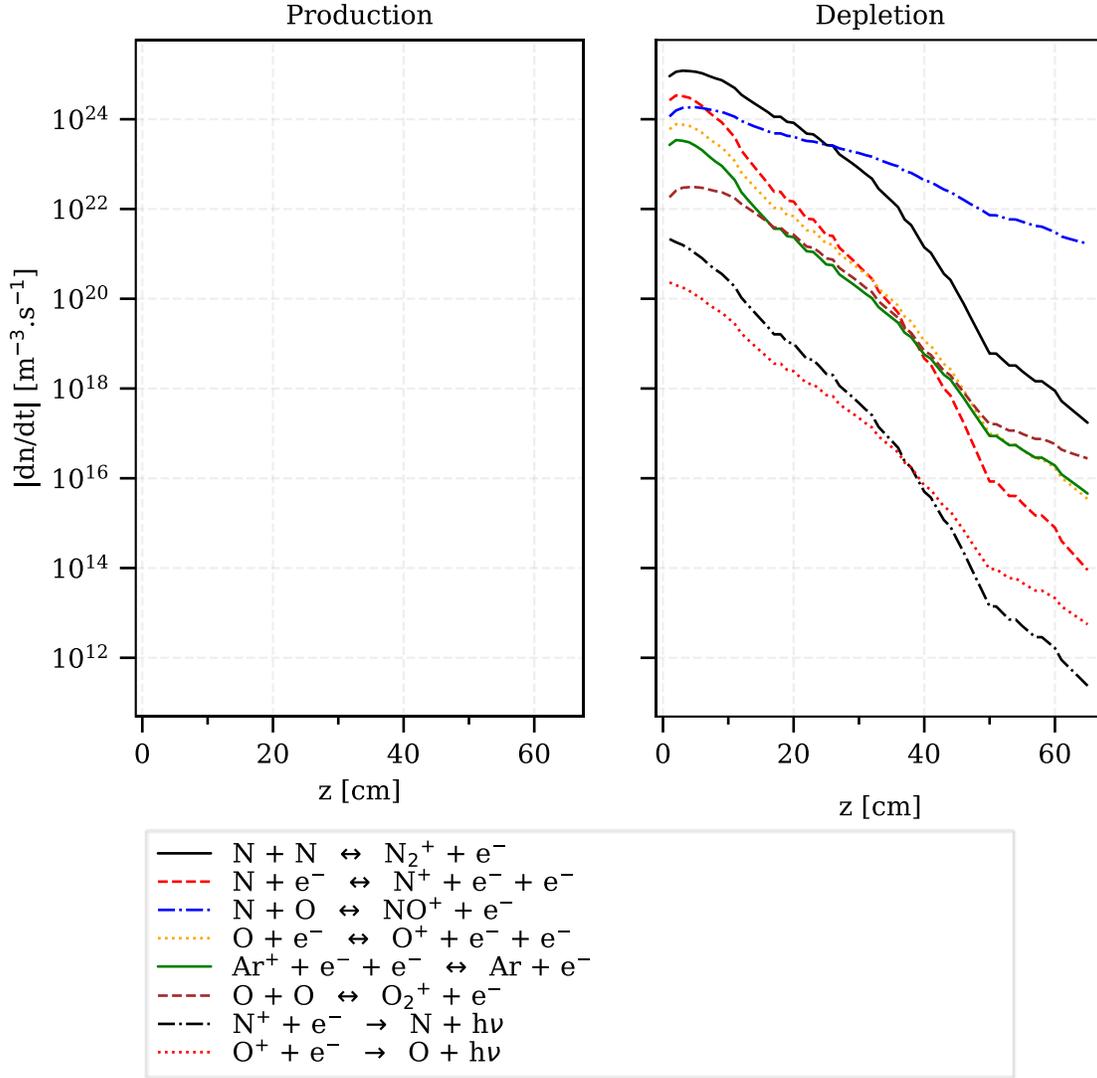
Figure 6 presents the electron number densities computed and measured in the water-cooled tube. For the sake of clarity, the error bars have only been indicated at the tube lengths where measurements have been made. The lines correspond to the average value of the simulations at a given tube length. This will be done for all the following figures presenting simulation results.

All models overpredict the electron number density, especially for longer tube lengths (above 40 cm). Between 10 and 15 cm, Johnston's and Cruden's models give electron densities between a factor 1.5 and 3 times higher than the measured electron densities, while our model agrees within a factor 2 with the measured densities. The largest discrepancies are observed at 40 cm, where all models predict densities at least one order of magnitude too high (from 11 times too high with our model up to 30 times with Cruden's model). The discrepancies are reduced at 50 and 65 cm: Cruden's model overpredicts the electron number density by a factor 15 and 20, Johnston's by a factor 4 to 7, and our model by a factor 2 to 4.5.



**Figure 6 Electron number densities along the tube: computations (lines), measurements [4,12] (black dots) and equilibrium (black diamonds).**

To gain insight on the simulation results, Figure 7 presents the net production and depletion of electrons along the tube length, for each reaction involving electrons. The dissociative recombination of  $\text{NO}^+$  and  $\text{N}_2^+$  are the main sources of electron depletion, especially for the longer tube lengths. After 30 cm, the dissociative recombination of  $\text{NO}^+$  becomes preponderant. But is it equilibrated?



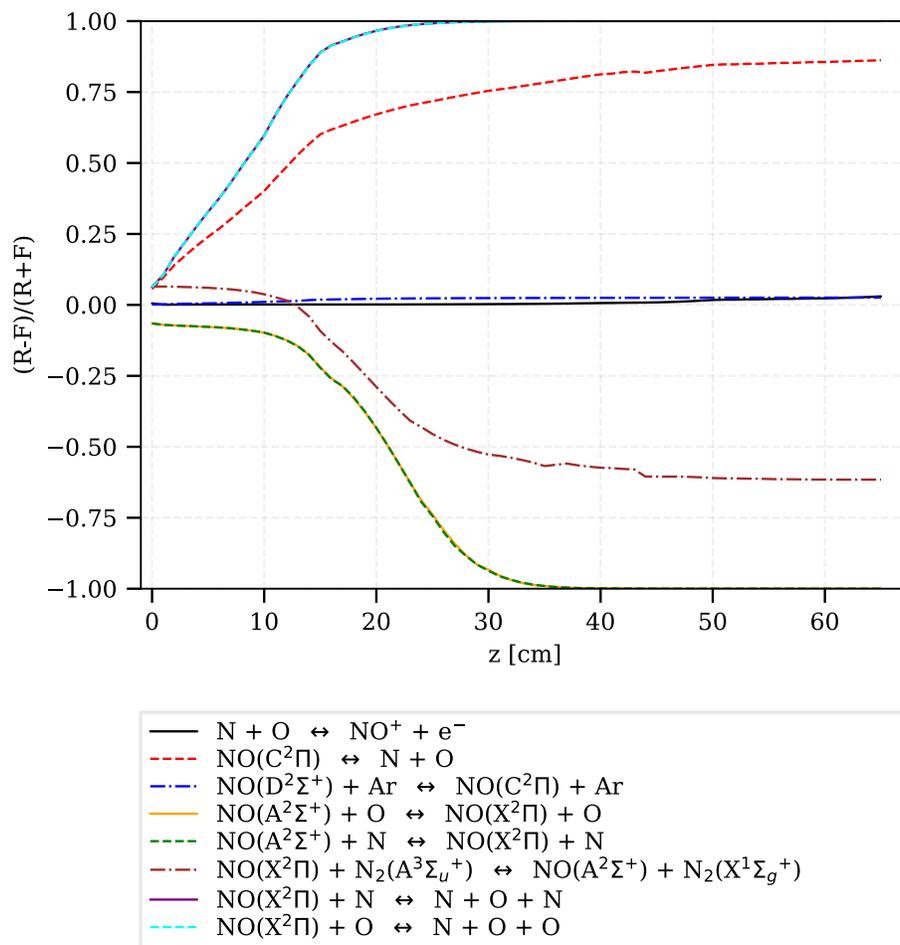
**Figure 7 Net rates of main reactions involving electrons. Rates are computed with the present model and plotted along the tube centerline.**

To know if a reaction is equilibrated, or if a certain direction (forward or reverse) is preponderant, one can look at the ratio of the reverse (R) rate minus the forward (F) rate over the sum of both rates:  $(R - F)/(R + F)$ . If this ratio is close to 0, then the reaction is equilibrated. If this ratio is close to 1, then the reverse reaction is preponderant ; if the ratio is close to -1, then the forward reaction is preponderant.

Figure 8 shows the evolution of this ratio for a selection of reactions along the tube length. Most reactions either tend to stay equilibrated along the tube axis, or tend to an extreme or another (ratios of 1 or -1). The dissociative recombination  $\text{NO}^+ + e^- \rightarrow \text{N} + \text{O}$  is equilibrated along the tube length. This can also be seen on Figure 11, with the superposition of the products of the overpopulation factors  $\rho_{\text{NO}^+}\rho_{e^-}$  and  $\rho_{\text{N}}\rho_{\text{O}}$  along the tube length. The main reactions that deplete N and O atoms in the mixture are reactions (7-8) given below:



These reactions are strongly non-equilibrated: their ratio  $(R-F)/(R+F)$  is close to 1 after 15 cm, meaning that they produce NO by consuming atomic nitrogen and oxygen. The ions and electrons are in partial equilibrium, and some charges are consumed by the dissociative recombination reaction  $\text{NO}^+ + e^- \rightarrow \text{N} + \text{O}$ , meaning that N and O atoms are in partial equilibrium with the electrons and  $\text{NO}^+$  ions, i.e. with the charged particles in the mixture. Hence, the disappearance of charges is rate-limited by the disappearance of N and O atoms through recombination of Nitric Oxide, with the reaction  $\text{N} + \text{O} + \text{M} \rightarrow \text{NO} + \text{M}$ . This explains the overpopulation of electrons in the water-cooled tube.

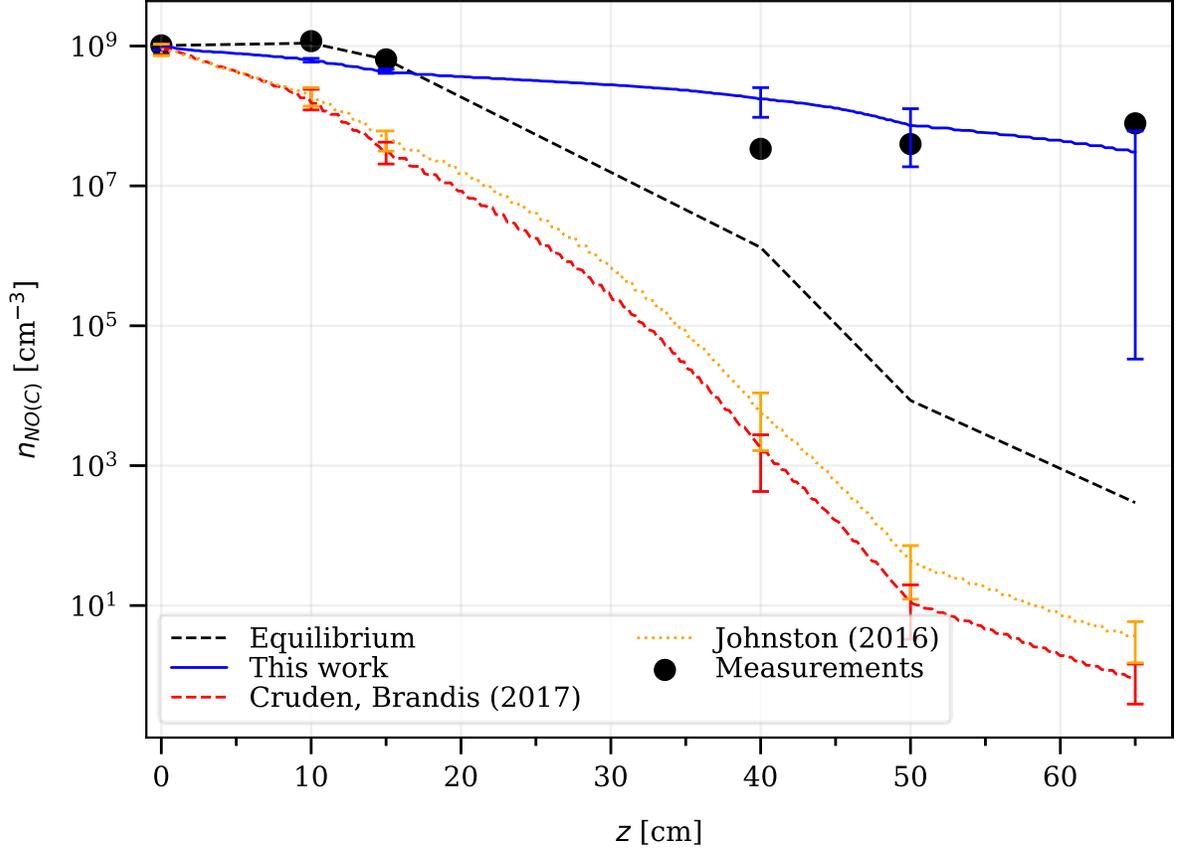


**Figure 8 Ratio of reverse (R) minus forward (F) rates over the sum of forward and reverse rates for a selection of reactions. A ratio of 0 means that the reaction is equilibrated.**

Figure 9 compares the densities of  $\text{NO}(\text{C}^2\Pi)$  computed with the different models to the measured and equilibrium densities in the water-cooled tube. All models relax towards equilibrium densities in approximately 250 ms for a sudden temperature drop, to be compared to the 1 to 2 ms of the residence time inside the water-cooled tube.

Both Cruden's and Johnston's models largely underpredict  $\text{NO}(\text{C}^2\Pi)$  densities in the water-cooled tube. Due to strong quenching of NO excited states,  $\text{NO}(\text{C}^2\Pi)$  is in partial equilibrium with  $\text{NO}(\text{X}^2\Pi)$ , which is largely dissociated at first, and recombines too slowly compared to the residence time inside the water-cooled tube.

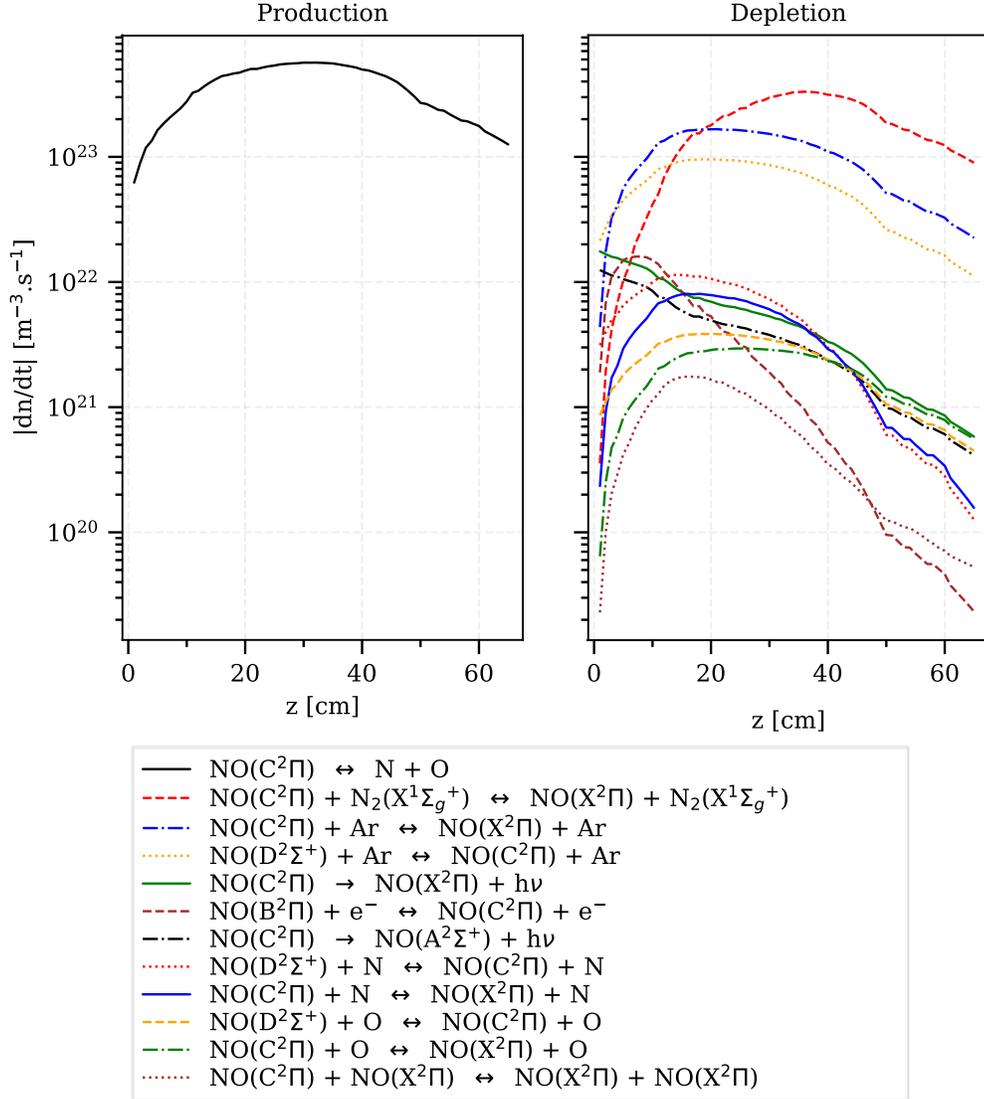
Our model, on the contrary, agrees well with the measured densities: between 0 and 15 cm, the model is within 50% discrepancies with the measurements. The largest discrepancies are at 40 cm, as for the electron density. The model overpredicts there  $\text{NO}(\text{C}^2\Pi)$  by a factor 4. These discrepancies are reduced to 25% (resp. 80%) at 50 cm (resp. 65 cm). The model converges to the equilibrium value in approximately the same time as Johnston's and Cruden's models.



**Figure 9** Densities of  $\text{NO}(\text{C}^2\Pi)$  along the tube centerline: computation (lines), measurements [4,12] (black dots) and equilibrium (black diamonds).

To understand the main processes responsible for the  $\text{NO}(\text{C}^2\Pi)$  population, we show on Figure 10 the main reactions populating or depleting  $\text{NO}(\text{C}^2\Pi)$  in the water-cooled tube.  $\text{NO}(\text{C}^2\Pi)$  is only populated through inverse predissociation. It is depleted through quenching, mainly by  $\text{N}_2$  and Argon, the main species present in the mixture. The partial equilibrium between  $\text{NO}(\text{C}^2\Pi)$  and  $\text{NO}(\text{X}^2\Pi)$  is thus challenged by the inverse predissociation, which tries to establish partial equilibrium between  $\text{NO}(\text{C}^2\Pi)$  and N and O atoms. This hypothesis of partial equilibrium between  $\text{NO}(\text{C}^2\Pi)$  and N and O atoms in the tube was first formulated by Gessman *et al.* [4].

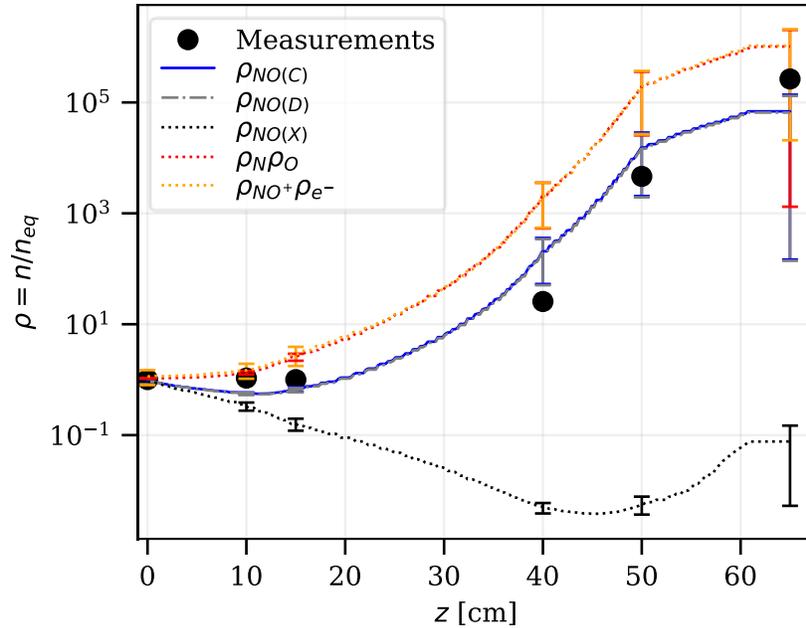
Also, an important process depleting  $\text{NO}(\text{C}^2\Pi)$  is the excitation to  $\text{NO}(\text{D}^2\Sigma^+)$  by collisions with Argon (the main species present in the mixture). This tends to put  $\text{NO}(\text{C}^2\Pi)$  and  $\text{NO}(\text{D}^2\Sigma^+)$  in partial equilibrium, which has been observed by Gessman *et al.* [4] in the measured spectra: both  $\text{NO}(\text{C}^2\Pi)$  and  $\text{NO}(\text{D}^2\Sigma^+)$  have the same measured overpopulation factors. Our model agrees: both  $\text{NO}(\text{C}^2\Pi)$  and  $\text{NO}(\text{D}^2\Sigma^+)$  have the same computed overpopulation factors, as shown in Figure 11 where the lines corresponding to these two electronic states are superposed, and the excitation reaction from  $\text{NO}(\text{C}^2\Pi)$  to  $\text{NO}(\text{D}^2\Sigma^+)$  is in equilibrium, as shown on Figure 8.



**Figure 10 Net rates of main reactions involving  $\text{NO}(\text{C}^2\Pi)$ . Rates are computed with the present model and plotted along the tube centerline.**

Figure 11 presents the computed overpopulation factors of different NO electronic states (X, C and D), as well as the measured overpopulation factors of  $\text{NO}(\text{C}^2\Pi)$  and the products of computed overpopulation factors of N and O and of ions  $\text{NO}^+$  and electrons. All computed overpopulation factors were obtained using the model described in this work. This figure allows to investigate the Gessman's hypothesis that  $\text{NO}(\text{C}^2\Pi)$  is in partial equilibrium with N and O and not with  $\text{NO}(\text{X}^2\Pi)$ . The important discrepancies between  $\text{NO}(\text{C}^2\Pi)$  and  $\text{NO}(\text{X}^2\Pi)$  overpopulation factors justify the second part of the hypothesis: the C and X states are not in partial equilibrium, especially after 10 cm. However, even if  $\rho_{\text{NO}(\text{C})}$  and  $\rho_{\text{NO}(\text{X})}$  follow the same trends, the existing discrepancies do not allow us to completely validate this hypothesis. It is possible to estimate  $\text{NO}(\text{C}^2\Pi)$  overpopulation using N and O overpopulations, but one might overpredict  $\text{NO}(\text{C}^2\Pi)$  overpopulation.

Moreover,  $\rho_{\text{N}}\rho_{\text{O}} \approx \rho_{\text{NO}^+}\rho_{e^-}$  in the water-cooled tube (the red and orange dotted lines are superposed). This is another way of observing that the dissociative recombination reaction  $\text{NO}^+ + e^- \rightarrow \text{N} + \text{O}$  is at equilibrium, as shown on Figure 8 and discussed earlier.

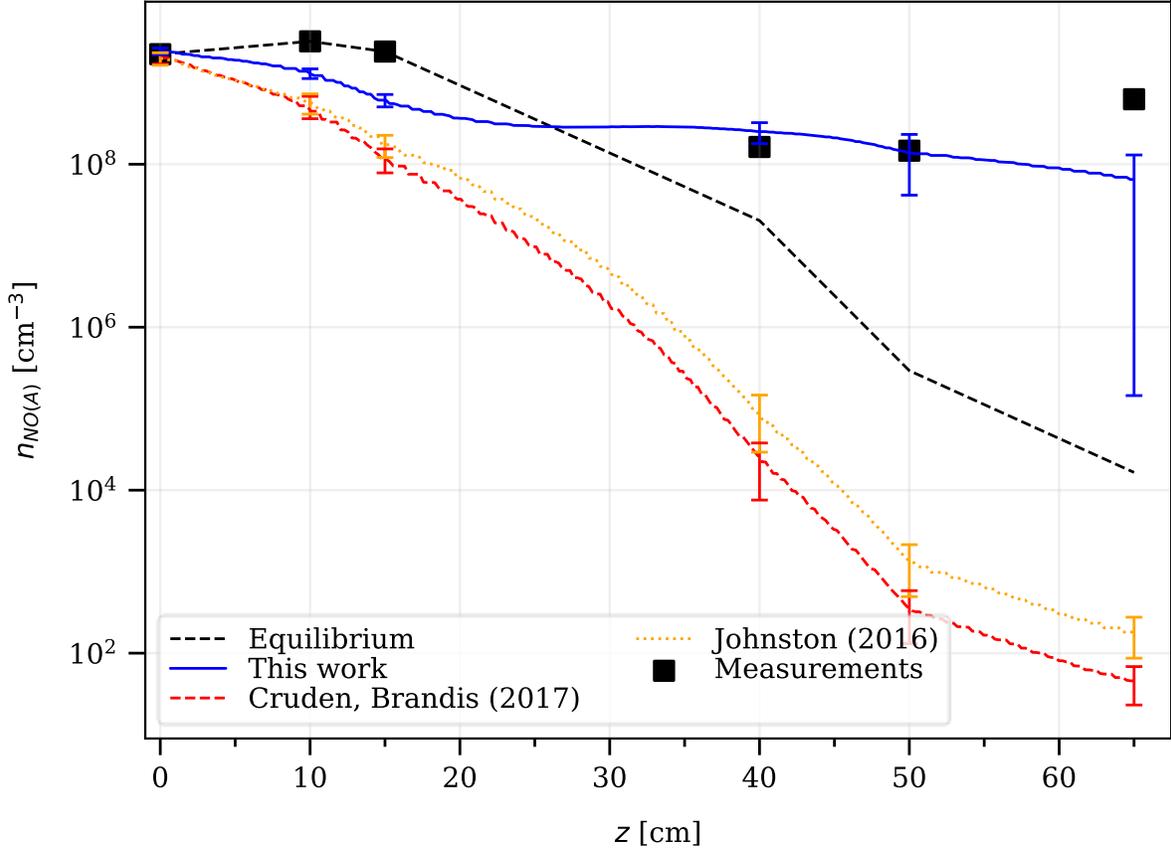


**Figure 11 NO( $C^2\Pi$ ) overpopulation factors along the tube centerline: computation (blue line) and measurements [4,12] (black circles). Computed overpopulations: NO( $D^2\Sigma^+$ ) (grey dash-dotted line), NO( $X^2\Pi$ ) (black dotted line). Products of computed overpopulations:  $\rho_{N\rho_O}$  (red dotted line) and  $\rho_{NO+\rho_{e^-}}$  (orange dotted line). Computation are performed with the present CR model.**

Figure 12 presents the computed densities of NO( $A^2\Sigma^+$ ) using the three models (Johnston's, Cruden's and ours), compared to the measured and equilibrium densities in the water-cooled tube.

The discrepancies between the predicted densities of NO( $A^2\Sigma^+$ ) and the measurements can be explained by the partial equilibrium between the A excited state and NO ground state due to important quenching reactions. NO, in its ground state, being largely dissociated at the entrance of the tube (due to the high temperature of approximately 8000 K), needs longer time to recombine than the residence time inside the water-cooled tube.

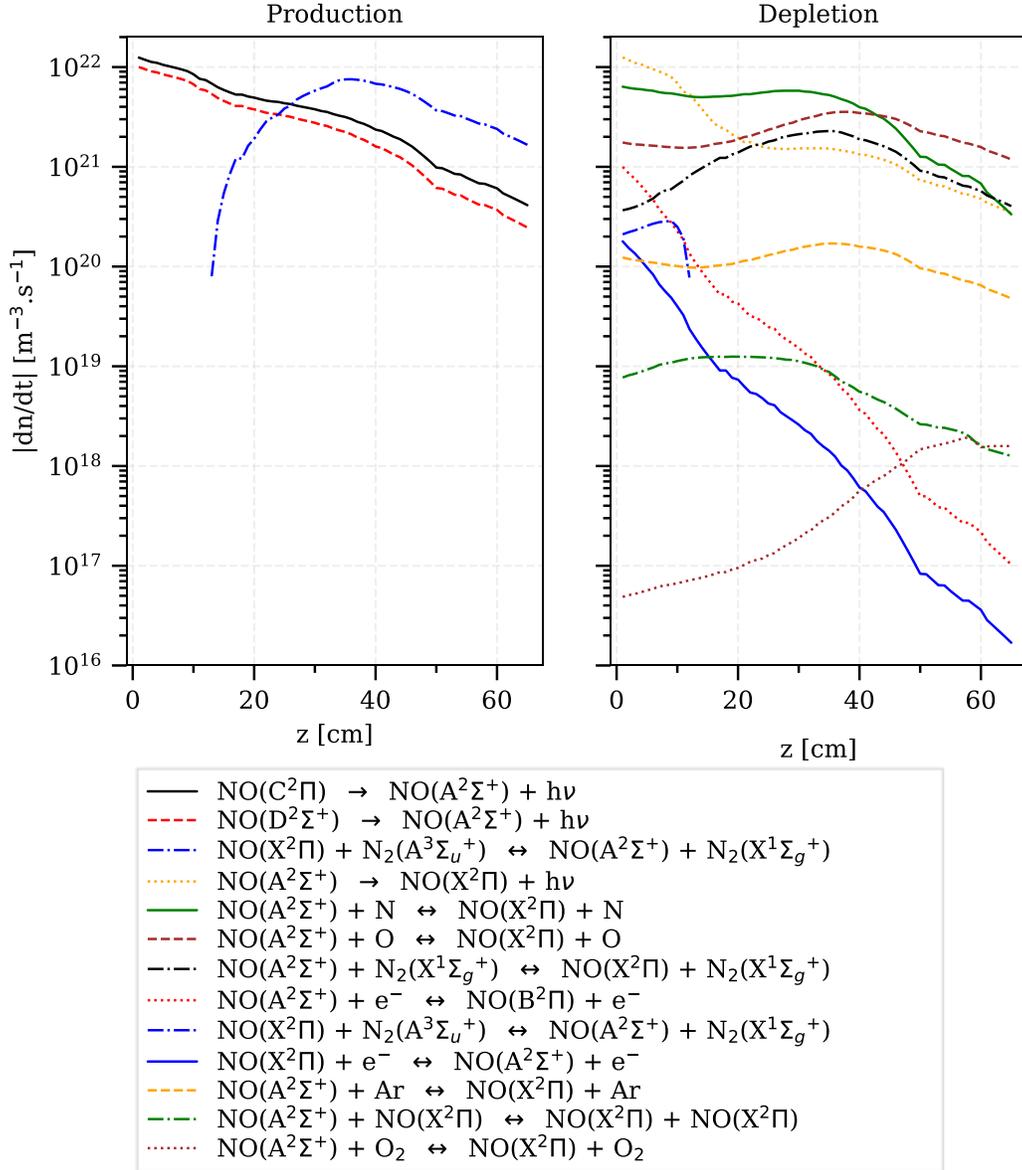
Our model, on the contrary, presents the same trends for NO( $A^2\Sigma^+$ ) densities than the measured ones. The mean values given by our model agree within 75% between 0 and 15 cm. These discrepancies are even reduced to approximately 30% at 40 and 50 cm, and stay between a factor 2 at 65 cm. NO( $A^2\Sigma^+$ ) density also converges to the equilibrium value outside the tube in approximately the same time as Johnston's and Cruden's models.



**Figure 12 Densities of  $\text{NO}(\text{A}^2\Sigma^+)$ : computation (lines), measurements [4,12] (black squares) and equilibrium (black diamonds).**

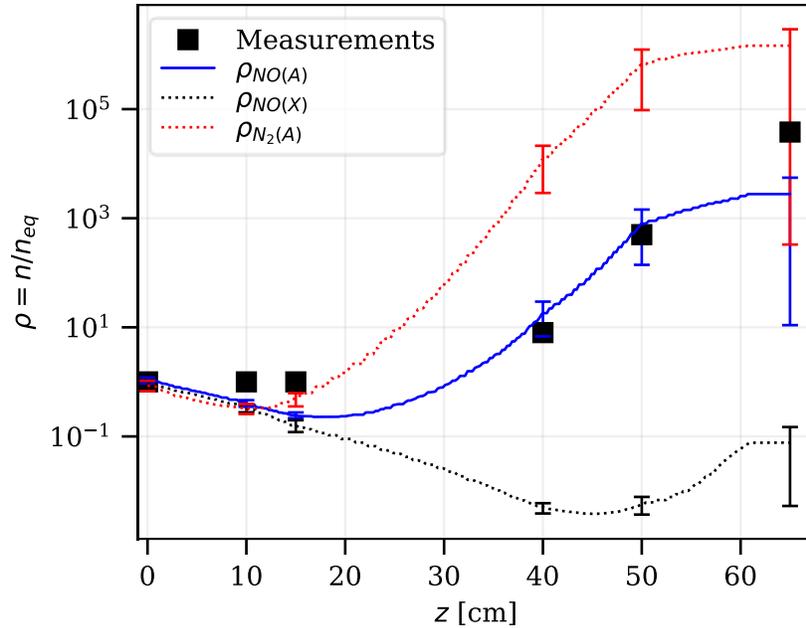
The main processes of  $\text{NO}(\text{A}^2\Sigma^+)$  population and depletion are presented on Figure 13. The A state of the NO molecule is mainly populated by spontaneous emission from the C and D states, and by collisions between  $\text{N}_2(\text{A}^3\Sigma_u^+)$  and  $\text{NO}(\text{X}^2\Pi)$ . Contrary to  $\text{NO}(\text{C}^2\Pi)$ , the spontaneous emission of  $\text{NO}(\text{A}^2\Sigma^+)$  is one of the main processes depleting the A state, alongside the quenching by N, O and  $\text{N}_2$ . The quenching by Argon depletes  $\text{NO}(\text{A}^2\Sigma^+)$  approximately one order of magnitude less than the quenching by  $\text{N}_2$ .

These processes compete to establish partial equilibrium between  $\text{NO}(\text{A}^2\Sigma^+)$  and other species. Firstly, the quenching reactions tend to establish partial equilibrium between  $\text{NO}(\text{A}^2\Sigma^+)$  and  $\text{NO}(\text{X}^2\Pi)$ . Secondly, the electronic energy transfer reaction tries to establish partial equilibrium between  $\text{NO}(\text{A}^2\Sigma^+)$  and  $\text{N}_2(\text{A}^3\Sigma_u^+)$ . Finally, the radiative decay of  $\text{NO} \text{C}^2\Pi$  and  $\text{D}^2\Sigma^+$  states toward  $\text{NO}(\text{A}^2\Sigma^+)$  disturbs these partial equilibrium by aligning  $\text{NO}(\text{A}^2\Sigma^+)$  overpopulation with the overpopulation of  $\text{NO} \text{C}^2\Pi$  and  $\text{D}^2\Sigma^+$  (which are equals, both states being in partial equilibrium as seen in Figure 8 and Figure 11).



**Figure 13 Net rates of main reactions involving  $\text{NO}(\text{A}^2\Sigma^+)$ . Rates are computed with the present model and plotted along the tube centerline.**

Figure 14 shows the overpopulation factors of  $\text{NO}(\text{A}^2\Sigma^+)$ , both computed with the model presented here and measured, and the computed overpopulation factors of  $\text{NO}(\text{X}^2\Pi)$  and  $\text{N}_2(\text{A}^3\Sigma_u^+)$  in the water-cooled tube. Between 0 and 15 cm,  $\text{NO}(\text{A}^2\Sigma^+)$  and  $\text{NO}(\text{X}^2\Pi)$  overpopulation factors are equal, meaning that the excited A state and the ground state are in partial equilibrium. After 40 cm, however, the excited A state is no longer in partial equilibrium with NO ground state. It is also not in partial equilibrium with  $\text{N}_2(\text{A}^3\Sigma_u^+)$ , even though both overpopulation factors seem to follow the same trends. Figure 8 also shows that the electronic energy transfer reaction between  $\text{N}_2(\text{A}^3\Sigma_u^+)$  and  $\text{NO}(\text{A}^2\Sigma^+)$  is not at equilibrium, and produces far more excited NO than excited  $\text{N}_2$ .



**Figure 14 NO( $A^2\Sigma^+$ ) overpopulation factors: computation (blue line), measurements [4,12] (black squares). Computed overpopulations factors: NO( $X^2\Pi$ ) (black dotted line) and  $N_2(A^3\Sigma_u^+)$  (red dotted line).**

More details on the impact of NO( $C^2\Pi$ ) predissociation and of the spontaneous emission of the C and D states toward NO( $A^2\Sigma^+$ ) are given in Section VIII (Appendix B). We summarize on Figure 15 the results given by the model described here. This figure is made of 2 different parts: for both of them, we present the results along the tube length (using  $z$ ), but also as a function of the measured temperature at various lengths and as a function of the time spent in the tube. This time given is an approximation obtained using Candler's temperature profile presented in Figure 5. We reproduce the overpopulation factors presented on Figure 11 and Figure 14. We also reproduce the simulated and measured densities of NO( $C^2\Pi$ ) and NO( $A^2\Sigma^+$ ). The errorbars on the densities at the exit of the tube ( $z = 65$  cm) go down to  $3.3 \cdot 10^4 \text{ cm}^{-3}$  for NO( $C^2\Pi$ ) and down to  $1.5 \cdot 10^5 \text{ cm}^{-3}$  for NO( $A^2\Sigma^+$ ).

The NO A and C excited states' overpopulation factors follow the same trend: this can be explained by the importance of NO( $C^2\Pi$ ) spontaneous emission toward NO( $A^2\Sigma^+$ ).

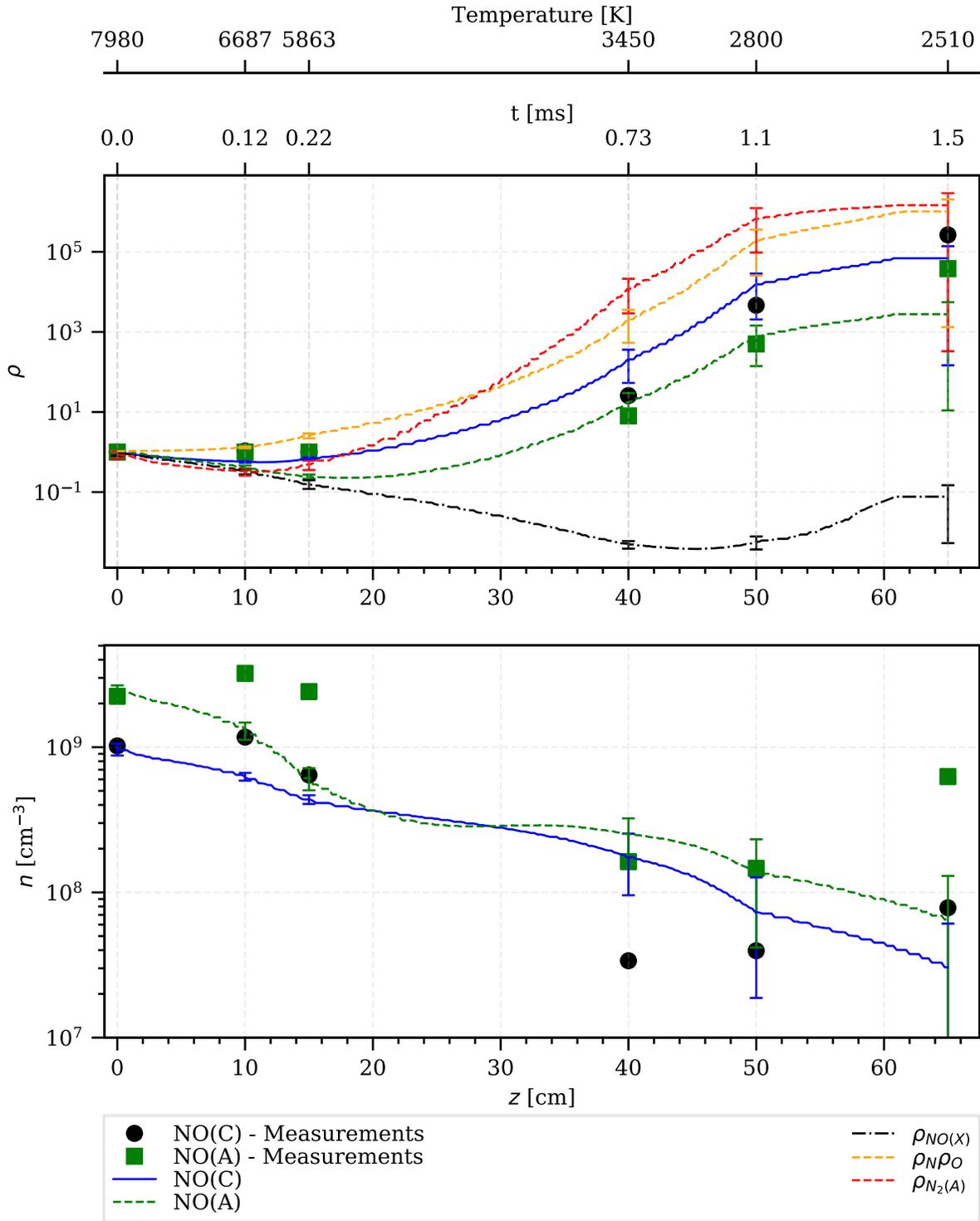


Figure 15 Summary of the results using the model presented in this work: (top) measured and computed overpopulation factors of NO(A $^2\Sigma^+$ ) (resp. green squares and green dotted line) and NO(C $^2\Pi$ ) (resp. black circles and blue line) compared to computed overpopulation factors of NO(X $^2\Pi$ ) (black dotted line),  $\rho_{N_2(O)}$  (orange dotted line) and  $\rho_{N_2(A)}$  (red dotted line); and (bottom) measured and computed densities of NO(A $^2\Sigma^+$ ) and NO(C $^2\Pi$ ), along the tube length and as function of time and temperature. Simulation results are obtained with the present model.

## V. Conclusions

We presented in this work an electronic-specific collisional-radiative model for Nitric Oxide, applied to an Air/Argon plasma produced by an ICP torch and undergoing fast recombination in a water-cooled tube. The present model, and the ones of Johnston [10] and Cruden and Brandis [11], were compared to measurements performed in the water-cooled tube by Gessman *et al.* [4,12].

Good agreement within a factor 4 (resp. 2) for the densities of the NO C<sup>2</sup>Π (resp. A<sup>2</sup>Σ<sup>+</sup>) state were obtained using the model presented in this work, while the models developed for shock-tube kinetics [10,11] differ by several orders of magnitude. This can be explained by NO(C<sup>2</sup>Π) formation via inverse predissociation and by the spontaneous emission of NO C and D states toward NO(A<sup>2</sup>Σ<sup>+</sup>), that were not considered in Johnston's or Cruden's models. Moreover, the quenching of NO excited states, especially with Argon and N<sub>2</sub>, is an important factor of depletion of NO A and NO C. Collisions with Argon allow the C and D states to be in partial equilibrium. The electronic energy transfer between N<sub>2</sub>(A<sup>3</sup>Σ<sup>+</sup>) and NO(A<sup>2</sup>Σ<sup>+</sup>) is also an important process populating the A state of NO.

Moreover, the charged species are in partial equilibrium with each other, owing to fast charge exchange and charge transfer. Furthermore, charged species are in partial equilibrium with the N and O atoms through the equilibrated dissociative recombination reaction NO<sup>+</sup> + e<sup>-</sup> ↔ N + O. Thus, the main path limiting electron recombination is the recombination of N and O atoms: N + O + M → NO + M. This slow process limits the depletion of charges in the mixture, and its rate is the main factor explaining the overpopulation of electrons in the water-cooled tube.

The present study gives a way to estimate the densities of the important radiative state NO(C<sup>2</sup>Π) from the densities of N, O and NO(X) without running an electronic specific CR model: the overpopulation of NO(C<sup>2</sup>Π) is approximately equal to the product of N and O overpopulations ( $\rho_{NO(C)} \approx \rho_N \rho_O$ ) because predissociation and its inverse processes are sufficiently fast to keep N, O and NO(C<sup>2</sup>Π) close to partial equilibrium. This equality holds within a factor 10 in the water-cooled tube. A more accurate prediction of NO(C<sup>2</sup>Π) densities requires the use of an electronic-specific model such as the one presented here.

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## VII. Appendix A

**Table 5 Model presented in this work**

Reaction	A [cm <sup>3</sup> /s]	n	E [K]	Ref.
N <sub>2</sub> + M → N + N + M				
M = N, O, N <sup>+</sup> , O <sup>+</sup>	4.98e-2	-1.60	113200	[9]
M = N <sub>2</sub>	4.27e4	-3.33	113200	[46]
M = Ar	4.98e-3	-1.60	113200	[13]
M = O <sub>2</sub> , N <sub>2</sub> <sup>+</sup> , O <sub>2</sub> <sup>+</sup> , NO <sup>**</sup> , NO <sup>+</sup>	1.16e-2	-1.60	113200	[9]
M = e <sup>-</sup>	1.99e1	-1.60	113200	[9]
O <sub>2</sub> + M → O + O + M				
M = N, O, N <sup>+</sup> , O <sup>+</sup>	1.66e-2	-1.50	59500	[9]
M = Ar	1.99e-10	0.00	54280	[47]
M = N <sub>2</sub> , O <sub>2</sub> , N <sub>2</sub> <sup>+</sup> , O <sub>2</sub> <sup>+</sup> , NO, NO <sup>+</sup>	3.32e-3	-1.50	59500	[9]
NO + M → N + O + M				
M = N, O, N <sup>+</sup> , O <sup>+</sup>	1.83e-7	0.00	75500	[9]
M = Ar	1.60e-9	0.00	74697	[48]
M = N <sub>2</sub> , O <sub>2</sub> , N <sub>2</sub> <sup>+</sup> , O <sub>2</sub> <sup>+</sup> , NO <sup>+</sup>	8.30e-9	0.00	75500	[9]
NO + O → N + O <sub>2</sub>	1.39e-11	0.00	19450	[9]
N <sub>2</sub> + O → NO + N	1.06e-6	-1.00	38400	[9]
N + O → NO <sup>+</sup> + e <sup>-</sup>	1.46e-15	1.00	31900	[9]
O + O → O <sub>2</sub> <sup>+</sup> + e <sup>-</sup>	1.18e-21	2.70	80600	[9]
N + N → N <sub>2</sub> <sup>+</sup> + e <sup>-</sup>	7.31e-17	1.50	67500	[9]
NO <sup>+</sup> + O → N <sup>+</sup> + O <sub>2</sub>	1.66e-12	0.50	77200	[9]
N <sup>+</sup> + N <sub>2</sub> → N <sub>2</sub> <sup>+</sup> + O <sub>2</sub>	1.66e-12	0.50	12200	[9]
O <sub>2</sub> <sup>+</sup> + N → N <sup>+</sup> + O <sub>2</sub>	1.44e-10	0.14	28600	[9]
O <sup>+</sup> + NO → N <sup>+</sup> + O <sub>2</sub>	2.32e-19	1.90	26600	[9]
O <sub>2</sub> <sup>+</sup> + N <sub>2</sub> → N <sub>2</sub> <sup>+</sup> + O <sub>2</sub>	1.64e-11	0.00	40700	[9]
O <sub>2</sub> <sup>+</sup> + O → O <sup>+</sup> + O <sub>2</sub>	6.64e-12	-0.09	18000	[9]

\*\* If not specified otherwise in the reaction formulas, NO molecules in this model are considered as NO(X<sup>2</sup>Π).

$\text{NO}^+ + \text{N} \rightarrow \text{O}^+ + \text{N}_2$	5.65e-11	-1.08	12800	[9]
$\text{NO}^+ + \text{O}_2 \rightarrow \text{O}_2^+ + \text{NO}$	3.99e-11	0.41	32600	[9]
$\text{NO}^+ + \text{O} \rightarrow \text{O}_2^+ + \text{N}$	1.20e-11	0.29	48600	[9]
$\text{O}^+ + \text{N}_2 \rightarrow \text{N}_2^+ + \text{O}$	1.51e-12	0.36	22800	[9]
$\text{NO}^+ + \text{N} \rightarrow \text{N}_2^+ + \text{O}$	1.20e-10	0.00	35500	[9]
$\text{O} + \text{e}^- \rightarrow \text{O}^+ + \text{e}^- + \text{e}^-$	6.48e9	-3.78	158500	[9]
$\text{N} + \text{e}^- \rightarrow \text{N}^+ + \text{e}^- + \text{e}^-$	4.15e10	-3.82	168600	[9]
$\text{O}^+ + \text{e}^- \rightarrow \text{O} + \text{h}\nu$	1.78e-13	-0.52	0	[9]
$\text{N}^+ + \text{e}^- \rightarrow \text{N} + \text{h}\nu$	2.52e-13	-0.48	0	[9]
$\text{Ar}^+ + \text{M} \rightarrow \text{Ar} + \text{M}^+$				
M = N <sub>2</sub> , O <sub>2</sub> , NO, N, O	1.49e-12	0.6	2260	[13]
$\text{Ar}^+ + \text{e}^- + \text{e}^- \rightarrow \text{Ar} + \text{e}^-$	2.90e-3	-1.0	-47800	[13]
$\text{H} + \text{e}^- \rightarrow \text{H}^+ + \text{e}^- + \text{e}^-$	2.51e7	-3.0	158000	[13]
$\text{H}^+ + \text{N} \rightarrow \text{H} + \text{N}^+$	3.32e-11	0.5	0	[13]
$\text{H}^+ + \text{O} \rightarrow \text{H} + \text{O}^+$	3.32e-11	0.5	0	[13]
$\text{NO(A)} + \text{Ar} \rightarrow \text{NO(X)} + \text{Ar}$	5.75e-15	0.5	0	This work
$\text{NO(A)} + \text{NO(X)} \rightarrow \text{NO(X)} + \text{NO(X)}$	1.44e-11	0.5	0	This work
$\text{NO(A)} + \text{O}_2 \rightarrow \text{NO(X)} + \text{O}_2$	8.63e-12	0.5	0	[11]
$\text{NO(A)} + \text{H}_2 \rightarrow \text{NO(X)} + \text{H}_2$	2.31e-13	0.5	0	This work
$\text{NO(A)} + \text{N}_2 \rightarrow \text{NO(X)} + \text{N}_2$	1.67e-20	2.74	182.8	This work
$\text{NO(B)} + \text{Ar} \rightarrow \text{NO(X)} + \text{Ar}$	5.77e-12	0.5	0	[11]
$\text{NO(B)} + \text{NO(X)} \rightarrow \text{NO(X)} + \text{NO(X)}$	1.15e-11	0.5	0	[11]
$\text{NO(B)} + \text{O}_2 \rightarrow \text{NO(X)} + \text{O}_2$	8.66e-12	0.5	0	[11]
$\text{NO(B)} + \text{N}_2 \rightarrow \text{NO(X)} + \text{N}_2$	3.52e-14	0.5	0	[11]
$\text{NO(C)} + \text{N}_2 \rightarrow \text{NO(X)} + \text{N}_2$	3.30e-10	0.5	0	This work
$\text{NO(C)} + \text{NO(X)} \rightarrow \text{NO(X)} + \text{NO(X)}$	1.30e-9	0.5	0	This work
$\text{NO(C)} + \text{O}_2 \rightarrow \text{NO(X)} + \text{O}_2$	8.66e-13	0.5	0	[11]
$\text{NO(C)} + \text{Ar} \rightarrow \text{NO(X)} + \text{Ar}$	5.77e-12	0.5	0	[11]
$\text{NO(D)} + \text{N}_2 \rightarrow \text{NO(X)} + \text{N}_2$	8.63e-12	0.5	0	This work
$\text{NO(D)} + \text{NO(X)} \rightarrow \text{NO(X)} + \text{NO(X)}$	2.09e-11	0.5	0	This work
$\text{NO(D)} + \text{O}_2 \rightarrow \text{NO(X)} + \text{O}_2$	3.49e-11	0.5	0	This work
$\text{NO(D)} + \text{Ar} \rightarrow \text{NO(X)} + \text{Ar}$	8.72e-12	0.5	0	This work
$\text{NO(D)} + \text{M} \rightarrow \text{NO(C)} + \text{M}$				
M = N, O, Ar	1.80e-10	0.5	0	[16]
$\text{NO(A,B,C,D)} + \text{M} \rightarrow \text{NO(X)} + \text{M}$	1.80e-10	0.5	0	This work
M = N, O				
$\text{NO(X)} + \text{e}^- \rightarrow \text{NO(A)} + \text{e}^-$	1.20e-10	0.0	63500	[14]
$\text{NO(X)} + \text{e}^- \rightarrow \text{NO(B)} + \text{e}^-$	1.01e-10	0.0	65700	[14]
$\text{NO(X)} + \text{e}^- \rightarrow \text{NO(C)} + \text{e}^-$	1.39e-10	0.0	74300	[14]
$\text{NO(X)} + \text{e}^- \rightarrow \text{NO(D)} + \text{e}^-$	1.39e-10	0.0	74300	[14]
$\text{NO(A)} + \text{e}^- \rightarrow \text{NO(B)} + \text{e}^-$	1.99e-9	0.0	0	[14]
$\text{NO(B)} + \text{e}^- \rightarrow \text{NO(C)} + \text{e}^-$	6.81e-9	0.0	0	[14]
$\text{NO(C)} \rightarrow \text{N} + \text{O}$	3.3e8 <sup>††</sup>	-	-	[16]
$\text{NO(A)} \rightarrow \text{NO(X)} + \text{h}\nu$	1.883e-7 <sup>‡‡</sup>	-	-	[16]
$\text{NO(B)} \rightarrow \text{NO(X)} + \text{h}\nu$	1.518e-6	-	-	[16]
$\text{NO(C)} \rightarrow \text{NO(X)} + \text{h}\nu$	5.251e-8	-	-	[16]
$\text{NO(C)} \rightarrow \text{NO(A)} + \text{h}\nu$	7.4e-8	-	-	[17]
$\text{NO(D)} \rightarrow \text{NO(X)} + \text{h}\nu$	2.725e-8	-	-	[16]
$\text{NO(D)} \rightarrow \text{NO(A)} + \text{h}\nu$	3.7e-8	-	-	This work
$\text{NO(B}^*) \rightarrow \text{NO(X)} + \text{h}\nu$	1.669e-7	-	-	[16]

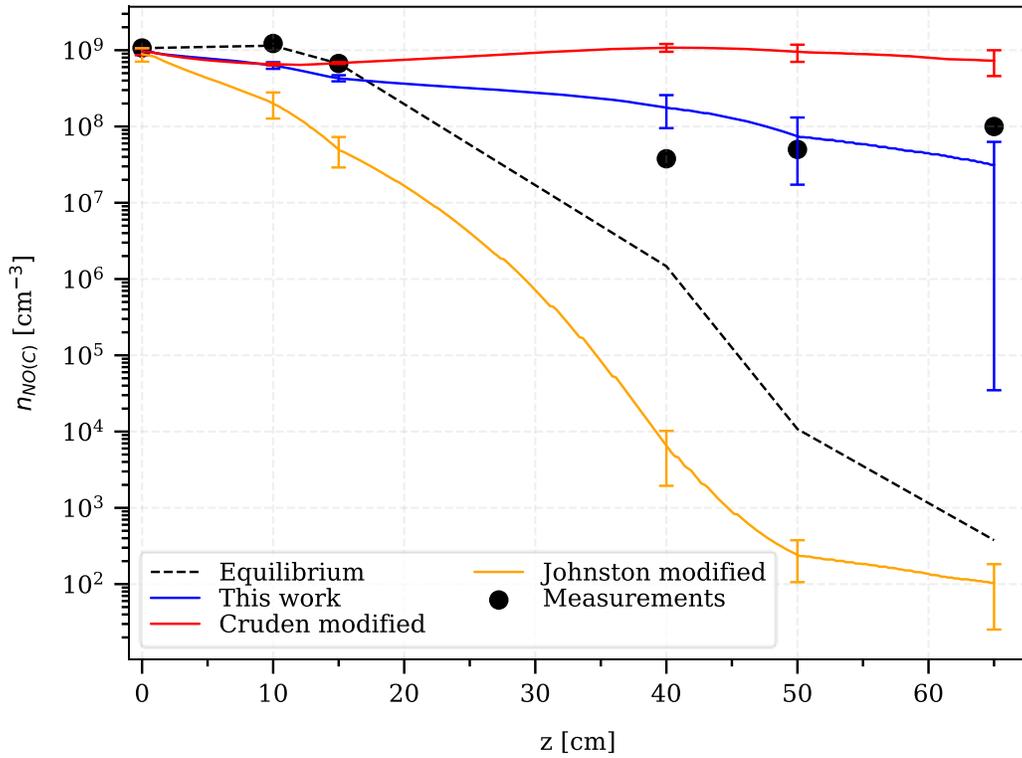
<sup>††</sup> In s<sup>-1</sup>. The rate constant of NO(C) predissociation depends heavily on the vibrational and rotational level (v,J): this A coefficient is between 3.3e7 and 2.0e10 s<sup>-1</sup>. Here we considered the value for the rovibrational states v=0, J>4.5 since approximately 96% of NO(C,v=0) levels have rotational numbers larger than 4.5 at 2300 K (which is approximately the temperature at the exit of the 65-cm tube).

<sup>‡‡</sup> Lifetimes in s.

$\text{NO(E)} \rightarrow \text{NO(X)} + \text{h}\nu$	$3.398\text{e-}8$	-	-	[16]
$\text{N}_2(\text{X})+\text{N}_2(\text{X})\rightarrow\text{N}_2(\text{A})+\text{N}_2(\text{X})$	$1.83\text{e-}12$	-0.5	71600	[14]
$\text{N}_2(\text{X})+\text{N}\rightarrow\text{N}_2(\text{A})+\text{N}$	$1.99\text{e-}7$	-1.5	71600	[14]
$\text{N}_2(\text{X})+\text{O}\rightarrow\text{N}_2(\text{A})+\text{O}$	$1.20\text{e-}7$	-1.5	71600	[14]
$\text{N}_2(\text{X})+\text{O}_2\rightarrow\text{N}_2(\text{A})+\text{O}_2$	$9.96\text{e-}9$	-1.5	71600	[14]
$\text{N}+\text{N}+\text{N}_2(\text{X})\rightarrow\text{N}_2(\text{A})+\text{N}_2(\text{X})$	$2.21\text{e-}31^{\S\S}$	-0.8	0	[14]
$\text{N} + \text{N} \rightarrow \text{N}_2(\text{A})$	$8.14\text{e-}17$	-0.35	0	[14]
$\text{N}_2(\text{A})+\text{N}_2(\text{A})\rightarrow\text{N}_2(\text{B})+\text{N}_2(\text{X})$	$1.23\text{e-}9$	0	0	[14]
$\text{N}_2(\text{X})+\text{e}^-\rightarrow\text{N}_2(\text{A})+\text{e}^-$	$1.02\text{e-}8$	0	71600	[14]
$\text{N}_2(\text{A})+\text{e}^-\rightarrow\text{N}_2(\text{B})+\text{e}^-$	$3.01\text{e-}10$	0	0	[14]
$\text{NO(X)}+\text{N}_2(\text{A})\rightarrow\text{NO(A)}+\text{N}_2(\text{X})$	$6.64\text{e-}11$	0	0	[35]

### VIII. Appendix B

To observe the impact of  $\text{NO(C}^2\text{II)}$  predissociation and of the spontaneous emission of the C and D states toward the A state, these reactions have been added to Johnston's and Cruden's models. The densities of the C and A states are shown respectively on Figure 16 and Figure 17. While it does not change fundamentally the behavior of Johnston's model, Cruden's model gives completely different results and is far closer to the measurements than before. The rate coefficients and radiative lifetime used are sufficient to compensate the quenching of the A and C states and to prevent these two excited states to be in partial equilibrium with NO ground state.



**Figure 16 Densities of  $\text{NO(C}^2\text{II)}$  computed (lines) with updated Johnston's and Cruden's models compared to measured (black dots) and equilibrium (black diamonds) densities.**

<sup>§§</sup> The unit for A is  $\text{cm}^6/\text{s}$ .

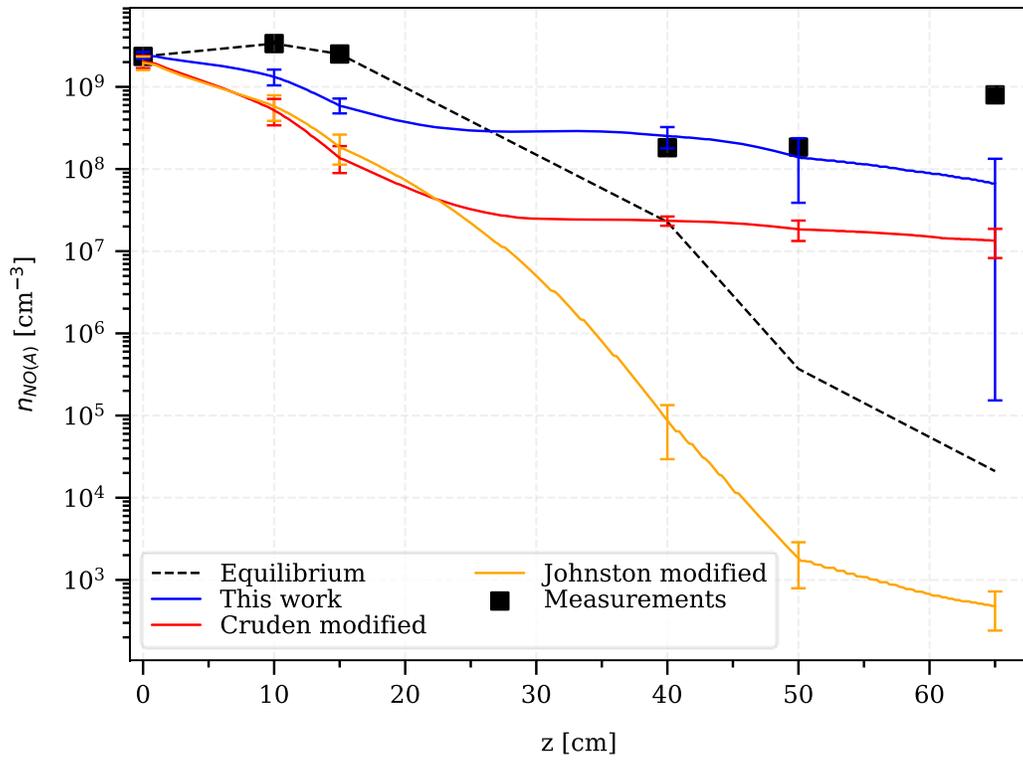


Figure 17 Densities of  $\text{NO}(A^2\Sigma^+)$  computed (lines) with updated Johnston's and Cruden's models compared to measured (black dots) and equilibrium (black diamonds) densities.