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Time-resolved *in situ* neutron diffraction study of $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite: a guide for the synthesis of complex chalcogenides

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ABSTRACT: Solid-state synthesis in sealed silica tube is one of the most popular methods for the production of inorganic materials. Nevertheless, the determination of the best synthesis conditions is a trial and error process that generally requires many experiments, varying temperature and duration time. In the present work, the synthesis of a polycrystalline sulfide by solid-liquid-gas reaction in sealed silica tube is, for the first time, monitored *in situ* by time-resolved neutron powder diffraction. All successive crystallization/decomposition sequences and intermediate products were analyzed in detail. These results were combined with *ex situ* X-ray powder diffraction and thermal analyses on crystallized samples. Our results revealed that sealed tube synthesis is highly dynamic with reaction times shorter than expected. Importantly, the experiments revealed the primordial role of the cooling conditions on the secondary phase formation and stability. Beyond new data on the complex crystal chemistry of germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$, our study demonstrates the compelling benefits of time-resolved *in situ* neutron powder diffraction in the optimization of the synthesis process for a wide range of complex chalcogenide materials.

INTRODUCTION

Solid-state synthesis in sealed tube is one of the fundamentals of inorganic material chemistry. Despite its simplicity, it presents several drawbacks: (i) the long duration (from hours up to days) that strongly depends on the temperature and the type of precursors, (ii) the difficulties in controlling chemical homogeneity, particle size and shape and (iii) the limitation to thermodynamically stable phases. In particular, the synthesis of ternary and quaternary chalcogenide mineral derivative compounds is quite challenging, as the synthesis and crystallization conditions in laboratory are far from the geological processes and durations. Typical synthesis process optimization consists in quenching the reaction at selected times and/or temperatures for *ex situ* characterization of the samples in order to determine the optimal conditions.¹⁻⁹ In addition to being time-consuming, such method is unsuitable for the detection and determination of quick reactions, as well as short-lived or air-sensitive intermediates. Therefore, the investigation of reaction pathways by *in situ* analysis has become an emerging field in solid-state chemistry.¹⁰⁻¹²

Real time *in situ* reaction monitoring is a powerful tool which allows to develop rational synthesis schemes and to better understand the reactions kinetic and mechanism

occurring during crystallization sequences.¹³⁻²² Such experiments have been mainly limited to XRD analyses.

Thanks to the development of dedicated instruments on high flux neutron sources enabling fast data acquisition,²³ *in situ* neutron powder diffraction has already been used to study the crystallization of amorphous materials,²⁴⁻²⁵ cationic ordering,²⁶⁻²⁹ mechanical properties,^{30,31} solid-gas reactions,³²⁻³⁵ and chemical processes in batteries.³⁶⁻³⁹ To this day, *in situ* powder neutron diffraction studies dedicated to sulfide materials have been performed only on pre-reacted or already formed sulfides,⁴⁰⁻⁴⁸ mainly to determine their temperature stability and decomposition mechanisms. In spite of the clear advantages of neutrons compared to X-rays (*i.e.* high penetration depth suited for bulk sample characterization, scattering lengths of the same order of magnitude even for light elements) and magnetic ordering detection, *in situ* neutron diffraction was never used, to the best of our knowledge, to characterize in real time the successive chemical reactions occurring during sealed tube synthesis of a sulfide from elemental precursors. Such experiments should be of prime interest for the determination and/or optimization of the synthesis conditions, especially in the case of complex compounds as encountered in quaternary sulfide systems.

In this study, we monitored for the first time the natures and crystal structures of the intermediate phases formed

during the solid-liquid-gas reactions of a quaternary sulfide, germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$. Using time-resolved *in situ* neutron powder diffraction data, we analyzed the successive chemical reactions as a function of temperature, and determined the key conditions to produce high purity samples. Such simple approach and methodology can provide guidance for the synthesis of many complex chalcogenides with the aim to design new efficient materials for various types of applications for energy harvesting and storage (photovoltaic, batteries, thermoelectric...).

EXPERIMENTAL SECTION

Synthesis. Cu (99 at.%, Alfa Aesar), Fe (99.5 at.%, Alfa Aesar), Ge (99.999 at.%, Alfa Aesar) and S (99.5 at.%, Alfa Aesar), all commercial powders, were stored and manipulated in a glove box under argon atmosphere. The powders were weighed in a stoichiometric ratio (22Cu: 8Fe: 4Ge: 32S) and ground together in an agate mortar. The 1.5 g batches of powder were then pressed into 5 mm diameter pellets. The pellets were placed in silica tubes ($\phi_{\text{ext}} = 8$ mm, $\phi_{\text{int}} = 6$ mm) evacuated down to a pressure of $\approx 10^{-2}$ mbar from an argon atmosphere. The tubes were sealed as close as possible to the sample to maximize the sample/silica weight ratio. Those tubes were used for *in situ* time resolved neutron powder diffraction.

In addition, the optimized synthesis conditions of $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite sample, determined from neutron powder diffraction analyses, were reproduced in laboratory. Sealed tubes were placed in a tubular furnace in vertical position with a heating rate of 2 K min^{-1} and a plateau at 973 K for 24 h. The samples were cooled down to 773 K at a natural cooling rate by switching off the heat power of the furnace, and then air quenched.

Differential scanning calorimetry and thermogravimetric analysis. Phase transitions and weight losses were investigated by differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA) respectively, using a Netzsch 404 F3 Pegasus apparatus. TGA-DSC measurements were carried out on a powder sample of reacted germanite from room temperature (RT) up to 973 K at the rate of 2 K min^{-1} under argon flow to avoid oxidation of the sample. The raw data were not corrected for instrumental contributions.

X-ray and neutron powder diffraction. The purity and the crystal structure of the germanite samples prepared in laboratory were checked by X-ray powder diffraction (XRPD). High-resolution data were collected using a Bruker D8 Advance Vario 1 fitted with a two-circle diffractometer (θ - 2θ Bragg-Brentano mode) equipped with a $\text{Cu K}\alpha$, X-ray tube, a Ge [111] monochromator and a LynxEye detector. Thermal stability investigations were carried out by XRD on a germanite powder sample obtained by sealed tube (ST) synthesis. Measurements were carried out between RT and 1000 K on powder in a glass capillary sealed under vacuum. The diffractograms were collected every 50 K during 2 h isotherms (4x30 min) using a Bruker D8 advance 2 diffractometer in Debye-Scherrer configuration equipped with a $\text{Mo K}\alpha$ source, a focusing multilayer mirror, a set of radial Soller slits and a LynxEye XE detector.

Time resolved *in situ* neutron powder diffraction (NPD) experiments were performed at the Institut Laue Langevin (ILL, Grenoble, France). The high temperature data were acquired on the high intensity two-axis powder D1B diffractometer, equipped with a vanadium furnace and a ^3He multi detector (1280 cells) covering a 128° angular range in a Debye-Scherrer geometry. Sealed silica tubes containing the elemental precursors were irradiated by the neutron beam ($\lambda = 1.284(1) \text{ \AA}$, refined using a standard compound), equipped with a Ge [311] monochromator. A first set of neutron powder diffraction (NPD) data was collected every 2.5 minutes from RT up to 973 K at a heating rate of 2 K min^{-1} (one pattern every 5 K), with 1 h isotherm at 675 K and 1.5 h isotherm at 973 K, followed by a fast cooling down to RT by furnace-quenching (switching off the power). A second set of NPD data was collected every 5 minutes from RT up to 973 K at a heating rate of 2 K min^{-1} (one pattern every 10 K), with 1 h isotherm at 973 K, and from 973 K down to RT at a cooling rate of 2 K min^{-1} (one pattern every 10 K). The temperature profiles are displayed in **Figures S1** and **S2**, respectively. The 1 h isotherm at 675 K in the former experiment was applied to limit the risk of explosion of the sealed silica ampoule, sometimes observed if uncontrolled exothermic reaction occurs between sulfur and metallic elements. From this experiment, it was shown that sulfur reacts with metallic elements well below 675 K (see below), consequently, this isotherm was not considered in the second *in situ* NPD experiment. Note that, with the second temperature profile, the sample temperature deviates from the set temperature below 580 K due to the furnace inertia. Considering the distance between the thermocouple and the sample and the duration of data acquisition (2.5 K min^{-1} or 5 K min^{-1}), the estimated error on the temperature is about $\pm 10 \text{ K}$. Refinements of the diffraction patterns (XRPD and NPD) were performed by Le Bail and Rietveld methods using the FullProf and WinPlotr software packages,^{49,50} with a manually selected background (mandatory due to the noisy background from silica contribution).

RESULTS AND DISCUSSION

Time-resolved *in situ* neutron diffraction. Heating. *In situ* NPD patterns were recorded every 5 K from RT up to 973 K (heating rate of 2 K min^{-1} , 2.5 min per pattern), with an isotherm of 1 h at 675 K. Before heating, a diffraction pattern was recorded during 1 h in order to determine the starting products (**Figure S3**). Rietveld refinement indicates the presence of the elementary components, Cu ($Fm\bar{3}m$, $a = 3.60(8) \text{ \AA}$), Fe ($Im\bar{3}m$, $a = 2.86(7) \text{ \AA}$), Ge ($Fd\bar{3}m$, $a = 5.64(7) \text{ \AA}$), α -S₈ ($Fddd$, $a = 10.41(7) \text{ \AA}$, $b = 12.89(4) \text{ \AA}$, $c = 24.35(1) \text{ \AA}$), and of CuS ($P6_3/mmc$, $a = 3.78(3) \text{ \AA}$, $c = 16.32(9) \text{ \AA}$). Note that CuS is likely formed during the sealing of the silica tube. Due to the semi-amorphous nature of the silica tube, all the diffraction patterns have a noisy background that hinders the possibility to reliably quantify the phases by Rietveld refinement.

The NPD patterns collected during the heating ramp from RT to 973 K are presented in **Figure 1** in a 2D contour plot between $20^\circ - 50^\circ$ in 2θ . The sequence of diffractograms could be divided in seven intervals characterized by the formation of the successive secondary, ternary and quaternary compounds. For the

sake of clarity, the phases are named according to their chemical formula although it does not necessarily represent their real final stoichiometry. The nature of each phase was determined by Rietveld refinement of the NPD patterns acquired at specific temperatures, *i.e.* the transition temperatures ($T = 300$ K, 430 K, 530 K, 590 K, 640 K, 770 K, 840 K and 973 K) and in-between ($T = 480$ K, 575 K, 615 K, 675 K and 810 K).

In the first temperature region, from 300 K to 430 K, the diffraction peaks of α -S₈ disappear close to its melting temperature (not observable on Figure 1 due to weak intensities) indicating the solid to liquid transition of the element, while Cu, Fe, Ge and CuS peaks remain unchanged. Above 430 K, the second region is characterized by the appearance of covellite CuS, concomitantly with the disappearance of Cu diffraction peaks suggesting the reaction of liquid sulfur with metallic copper (Figure S4). At 490 K, a reaction between sulfur and iron occurs leading to the decrease of the peak intensity of Fe and the appearance of FeS ($P6_3mc$, $a \sim 7.0$ Å, $c \sim 5.6$ Å) diffraction peaks. Starting from 530 K, the third temperature interval is characterized by the formation of CuFeS₂ chalcopyrite ($I\bar{4}2d$) (Figure S5). Concomitantly, the intensity of the diffraction peaks of Fe strongly decreases up to 590 K, while it slightly decreases for CuS. Considering that the peaks intensity of the FeS phase remains constant from 530 K to 590 K, it suggests that Fe reacts with CuS to form CuFeS₂ chalcopyrite ($I\bar{4}2d$, $a \sim 5.2$ Å, $c \sim 10.4$ Å). Hence, at 590 K, Cu and Fe have totally reacted with S to form CuS, FeS and CuFeS₂ binary and ternary phases, while Ge remains in its elementary state (Figure S5). Above 590 K, the contents of CuS, FeS and CuFeS₂ decrease up to 640 K, progressively replaced by pyrite FeS₂ ($Pa\bar{3}$, $a \sim 5.4$ Å), nukundamite Cu_{0.85}Fe_{0.15}S ($P\bar{3}m1$, $a \sim 3.8$ Å, $c \sim 11.2$ Å), and bornite Cu₅FeS₄ ($Fm\bar{3}m$, $a \sim 5.5$ Å) (Figure S6). These observations are consistent with the stability of chalcopyrite CuFeS₂ and pyrrhotite FeS up to 600 K.⁵¹ The formation of bornite Cu₅FeS₄ starts around ~ 575 K. Its identity was later confirmed by Rietveld

refinements (at higher temperature) and by its characteristic topotactic transition temperatures during the subsequent cooling ramp.⁵²⁻⁵⁴ No apparent variation of the peaks intensity was detected from 640 K up to 770 K (Figure 1).

Above 770 K, the diffraction peaks of nukundamite Cu_{0.85}Fe_{0.15}S disappear, while the peaks intensity of bornite Cu₅FeS₄ drastically increases (Figure S7), indicating the decomposition of nukundamite into a bornite phase probably associated with the formation of digenite Cu_{1.8}S. This observation agrees with data reported in the literature.⁵¹ Additionally, one can note the appearance at 810 K of an unidentified diffraction peak (marked by a question mark) concomitant with the slow fading of pyrite FeS₂ and Ge diffraction peaks, until its total disappearance between 840 K and 860 K (Figure S7 and Figure S8). Finally, the first quaternary phase (HT_quat) appears above 840 K from the reaction of Ge with the ternary phases, *e.g.* nukundamite and bornite (Figure S8). Experimentally, the heating ramp was followed by an isotherm of 90 min at 973 K, where the content of HT_quat phase increases slowly at the expense of bornite during the first 60 min (Figure S9). It indicates that the formation of the high temperature quaternary phase was completed during the isotherm.

From the present data, it is quite difficult to determine the crystal system of the high temperature quaternary (HT_quat) phase: it can be either a high temperature disordered cubic form of the expected germanite compound or an intermediate phase with lower symmetry and different unit cell parameters. From the position of the reflection peaks and their relative intensities, it is obvious that the HT_quat phase has a metric close to cubic germanite ($P\bar{4}3n$), but it is also true for many quaternary compounds. Unfortunately, the experimental setup prevents to discriminate between cubic and tetragonal cell, as the resolution is not sufficient to detect eventual peak splitting, representative of a tetragonal symmetry.

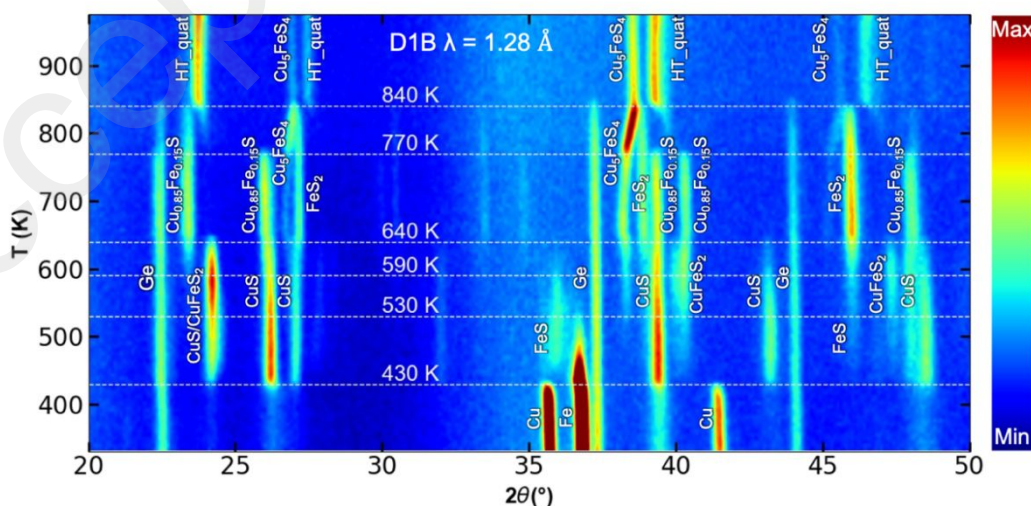


Figure 1. 2D contour plot showing the NPD patterns of the in situ sealed tube reaction from elemental precursors

Tetragonal $P\bar{4}2c$ space group, as found in renierite,⁵⁵ is a maximal *translationengleiche* subgroup of the cubic space group found in germanite ($P\bar{4}3n$). Yet, the experimental conditions did not allow to detect the presence (or absence) of the renierite structure. Thus, a briartite phase ($I\bar{4}2m$)⁵⁶ was used for the refinement of the HT_quat phase as it generates fewer reflections than renierite ($P\bar{4}2c$) and hence lowers the risk of an artificially low χ^2 factor. **Figure 2a** shows how the supplementary reflections allowed by the $P\bar{4}2c$ space group, which are extinguished in the calculated intensity profile of renierite, artificially improves the χ^2 factor from 1004 for $I\bar{4}2m$ to 700.9 for $P\bar{4}2c$ (**Table S1**). Additionally, a closer look at the peak shape reveals a wide Gaussian tail on the right of the reflection peaks 440 and 622 while the 222 reflection peak is not affected. It suggests that only reflections influenced by a cubic to tetragonal deformation have this peak shape anomaly. One origin of the peak broadening is the distribution of the cell parameters induced by microdeformations, which lead to microstrains. While isotropic microstrain produces a $\tan\theta$ dependent broadening, an additional hkl -dependence suggests that the microstrain is anisotropic. Although both isotropic and anisotropic strain models adequately describe the peak profile of the 222 reflection (**Figure 2b**), that of 440 (**Figure 2c**) was greatly improved by applying Stephens anisotropic microstrain model (pink).⁵⁷ The standard isotropic strain model (blue) slightly underestimates the width of the 440 peaks. In the case of an anisotropic strain of a cubic space group ($P\bar{4}3m$ red, $P\bar{4}3n$ yellow, **Figure 2d**), the broadening would be symmetrical on both sides of the peak, thus excluding this option. The parameters used for these refinements are gathered in **Table S1**.

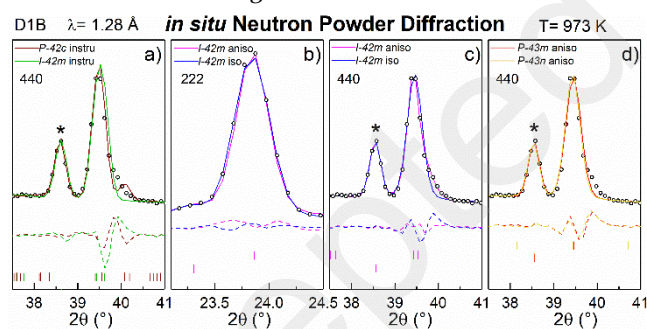


Figure 2. Different Le Bail fits of the NPD isotherm at 973 K of the HT_quat phase: a) $P\bar{4}2c$ (brown) and $I\bar{4}2m$ (green) refined with instrument resolution parameter only, b) and c) $I\bar{4}2m$ refined with an isotropic (blue) and anisotropic (pink) strain for reflections 222 and 440, respectively and d) $P\bar{4}3m$ (red) and $P\bar{4}3n$ (yellow) refined with an anisotropic strain. An asterisk pinpoints a contribution of the high temperature form of bornite Cu_5FeS_4 .

Stephens microstrains model largely captures the anisotropy of the broadening determined by individual fits and induces a clear improvement of the peak shape fitting. Furthermore, it would be interesting to verify the presence of such microdeformations by transmission electron microscopy, but it is technically impossible since this phase only exists at high temperature (above 860 K). To

yield the sample utmost closely related to the HT composition, a sealed tube was air quenched at 973 K in laboratory. The diffraction pattern is clearly not that of a pure cubic sample due to the splitting of the peaks at $\sim 34.0^\circ$, $\sim 48.5^\circ$ and $\sim 58.0^\circ$ (**Figure S10**). However, it is also not the diffraction pattern expected for a pure tetragonal sample as the relative intensities of the diffraction peaks deviate from the expected 2:1 ratio for a tetragonal symmetry, notably for the 440 and 622 reflections. Thus, the refinement of this diffraction pattern requires to consider a mixture of tetragonal and cubic phases with the same metric. The a/c ratio of the tetragonal phase is ~ 1.006 , which is expected for a renierite phase.⁵⁸ These results are in favor of the coexistence of both tetragonal and cubic phases at 973 K. Finally, the absence of superstructure reflections on both NPD pattern at 973 K (**Figure S9**) and XRPD pattern after quenching (**Figure S10**) can be explained by the quench at high temperature which may induce cationic disorder.

Cooling. The cooling conditions were investigated in real time by *in situ* neutron powder diffraction in two separate experiments: a quench at 973 K and a controlled cooling at the rate of 2 K min^{-1} . For both cooling pathways, the same main phase was obtained at RT (**Figures S11a** and **S11b**), the only difference results in the presence of Cu_5FeS_4 (bornite HT cubic form) along with an unknown phase in the quenched sample. The presence of superstructure peaks at low angle (**Figures S11a** and **S11b**, first inset on the left), and the absence of splitting of the reflection 800 (**Figure S11a** and **S11b**, inset on the right), indicate the formation of germanite and the absence of renierite, respectively.

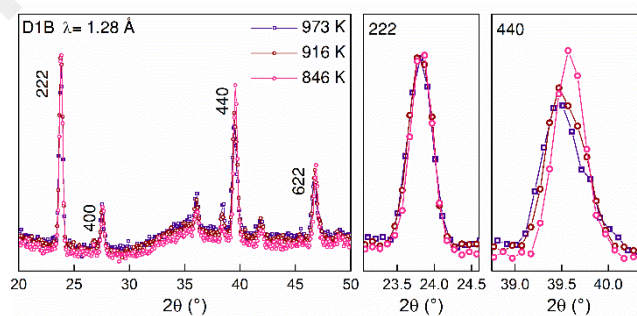


Figure 3. NPD patterns acquired during controlled cooling (2 K min^{-1}) at 973 K, 916 K and 846 K, showing the change from tetragonal to cubic structure during cooling.

The next paragraph describes the changes in the diffractograms measured during the controlled cooling (2 K min^{-1}). The noisy background of the patterns acquired during the cooling ramp (due to a short collection time and silica contribution) together with the shrinking cell parameters, both consequences of the NPD *in situ* experiment conditions, hinder the refinement of the anisotropic microstrain parameters. Thus, the structural changes during the cooling ramp will be described using a qualitative approach. During controlled cooling, the first observable change is related to the 440 peak profile, which evolves from a well-defined doublet at 973 K to a thinner single peak at 846 K passing through a broad “single” peak

at 916 K (**Figure 3**). This peak profile evolution suggests the appearance of a cubic phase, with possibly a different composition, besides the tetragonal one as the temperature decreases from 973 K to 846 K. The broadening of the “single” peak at 916 K can be attributed to cationic disorder and/or to the coexistence of both cubic and tetragonal phases.^{59,60} The second observable change is the emergence of a peak at 25.9° in the temperature range 860 K - 820 K. This peak can be attributed to the superstructure reflection 321 of germanite (**Figure S12**), indicating that the cubic phase formed between 916 K and 846 K is germanite. All superstructure peaks should be observed simultaneously, but the noisy background coupled to the intrinsically low intensity of the superstructure peaks make their detection rather difficult. The 321 is the first germanite superstructure peak detected due to its higher angle (25.8°), where the silica contribution to the background is lower. Finally, the third observable change is the anomalous broadening of the main peaks base, observed while cooling the sample below 770 K (**Figure S13**) and attributed to the appearance of a low crystallinity sphalerite derivative phase. The formation of this phase below 770 K limits the possibility to obtain a pure $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite sample by using a slow cooling down to room temperature despite the apparent high purity of the sample obtained using controlled cooling conditions (**Figure S14**).

Following these new insights on germanite crystallization and neighboring phases formation and stability, an optimized sealed tube synthesis route was proposed for $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite: heating at 973 K for 24 h, cooled down to 773 K at a natural cooling rate by switching off the power of the furnace and then air quenched. With these optimized conditions, a high purity germanite sample of 5 grams was prepared by sealed tube synthesis (**Figure 4**). Its thermoelectric properties after densification either by spark plasma sintering or hot

pressing were recently reported and compared to those of germanite sample prepared by mechanical alloying.⁶¹

Thermal stability of germanite. Considering the difficulty to prepare a high purity germanite sample, due to the easy formation and stability of neighboring phases, it is important to acknowledge the thermal stability of germanite as well as its reactivity zones and products of decomposition. For that matter, a TGA-DSC analysis and XRPD measurements in temperature were performed on a high purity sample of $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite prepared by optimized sealed tube synthesis.

On the TGA curve (**Figure 5**), slope changes are observed at 650 K and 870 K, both referring to an increase in sulfur loss (~ 0.5 wt. % at 650 K and ~ 3.1 wt. % at 870 K). The DSC curve exhibits a broad endothermic signal starting at ~ 870 K, which coincides with the second slope change on the TGA curve, suggesting an evolution of the germanite sample above this temperature.

The X-ray diffractograms, recorded from RT to 1000 K on powder sample in a glass capillary sealed under vacuum (2 h isotherms, every 50 K), are displayed in **Figure 6**. Due to presence of two spurious peak at ~ 7° and ~ 11° caused by the furnace window contribution (Al foil), the detection of the superstructure peaks was mainly limited to the 211 and 321 reflections. The thermodiffractograms exhibit two main events at ~ 650 K and at ~ 900 K (**Figure 6**). The first one, occurring around 650 K, is related to the increase of the diffraction peak at ~ 21° indicating the formation of bornite Cu_5FeS_4 phase in its high temperature cubic form. This secondary phase formation can be related with the first step of sulfur volatilization detected by TGA above 650 K (**Figure 5**). It indicates that the first step of sulfur volatilization during high temperature exposure of germanite leads to the formation of the copper-rich Cu_5FeS_4 phase (M_{tot}/S ratio of 1.5) instead of sulfur deficient germanite derivative (M_{tot}/S ratio ≤ 1.0625).

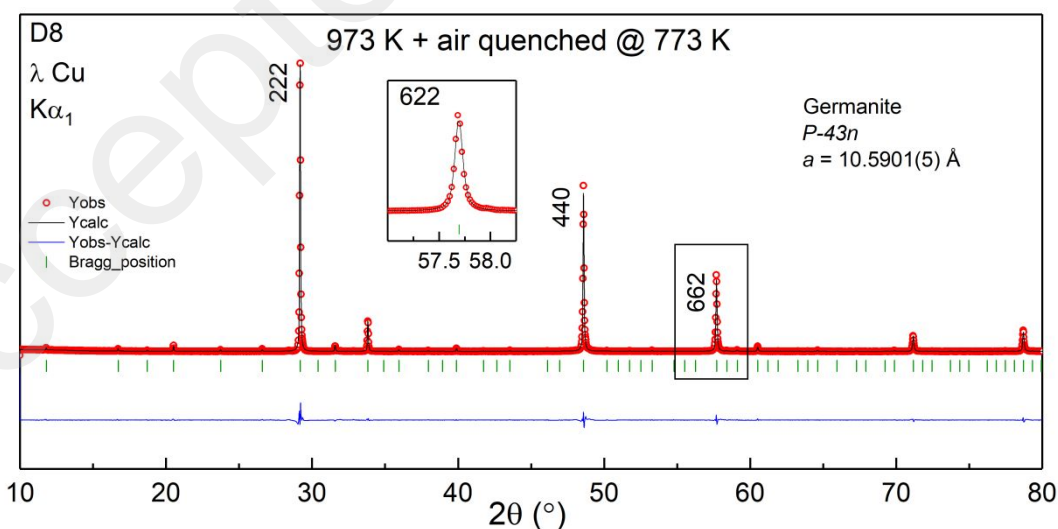


Figure 4. Refinement of the XRPD pattern measured at room temperature of a germanite sample prepared by optimized sealed tube synthesis.

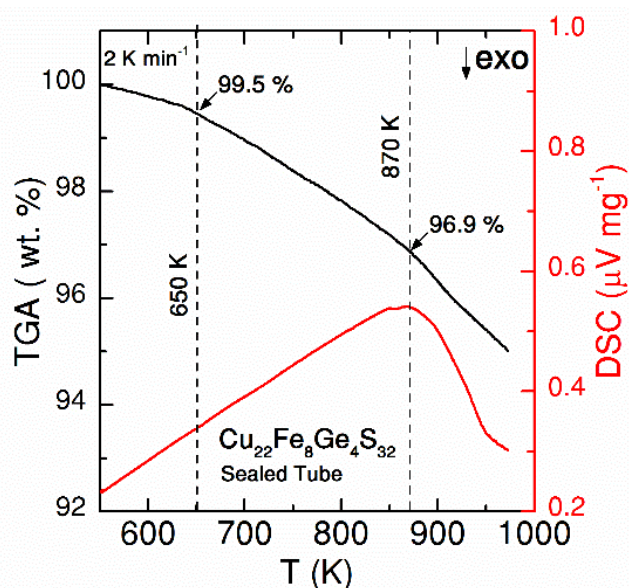


Figure 5. TGA-DSC curves of a germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ sample in a platinum pan.

The second modification, occurring around 900 K is related to the disappearance of the superstructure peaks 211 and 321 of germanite, concomitantly with the appearance of new peaks next to the 440 and 622 reflections. This pattern modification through the disappearance/appearance of diffraction peaks is the signature of the change of the cubic germanite structure to a closely related tetragonal structure as the temperature is increased, but with a different chemical composition. A similar structural evolution was previously observed for the Zn for Cu substitution in the germanite-renierite series $\text{Cu}_{22-x}\text{Zn}_x\text{Fe}_8\text{Ge}_4\text{S}_{32}$,⁶² where the cubic germanite transforms into the tetragonal renierite as x is increased. Indeed, such structural modification from cubic to tetragonal structure induces the splitting of equivalent reflections into two

reflections for the hhl and hkk reflections allowed by the cubic space group, e.g. 440 and 622, or into three reflections for the hkl reflections.⁶² However, in the present case, the formation of the tetragonal phase is not induced by direct element substitution, but is indirectly related to the sulfur loss, as detected by TGA above 870 K (Figure 5). Indeed, as shown in the first part of germanite sample evolution, sulfur volatilization induces the formation of the copper-rich phase Cu_5FeS_4 bornite, and consequently, it results in a modification of the germanite structure and composition due to the germanium enrichment. This germanium enrichment is most probably the reason of the high temperature phase change from cubic germanite to tetragonal renierite structure (Figure S14).

By considering that a chemical composition modification of germanite phase induced by sulfur volatilization at high temperature should modify significantly its unit cell parameter evolution, a second XRPD experiment was performed from room temperature up to 860 K, i.e. below the temperature of modification from cubic to tetragonal structure (Figure S15). The thermal evolution of the unit cell parameter of germanite obtained from refinement of the patterns by the Le Bail method is shown in Figure 7. The linear evolution of the unit cell of germanite up to 650-700 K demonstrates its good chemical stability in this temperature window. The steeper slope above 650 K, supported by sulfur volatilization above 650 K (Figure 5), suggests a chemical variation (i.e. cationic disordering and/or composition variation) within the crystal, as already shown in structurally related colusites.^{48,63-65} Moreover, this experiment shows that, at 860 K, the superstructure peaks of germanite were still present and that the cubic to tetragonal phase transformation is not yet achieved (Figure S15).

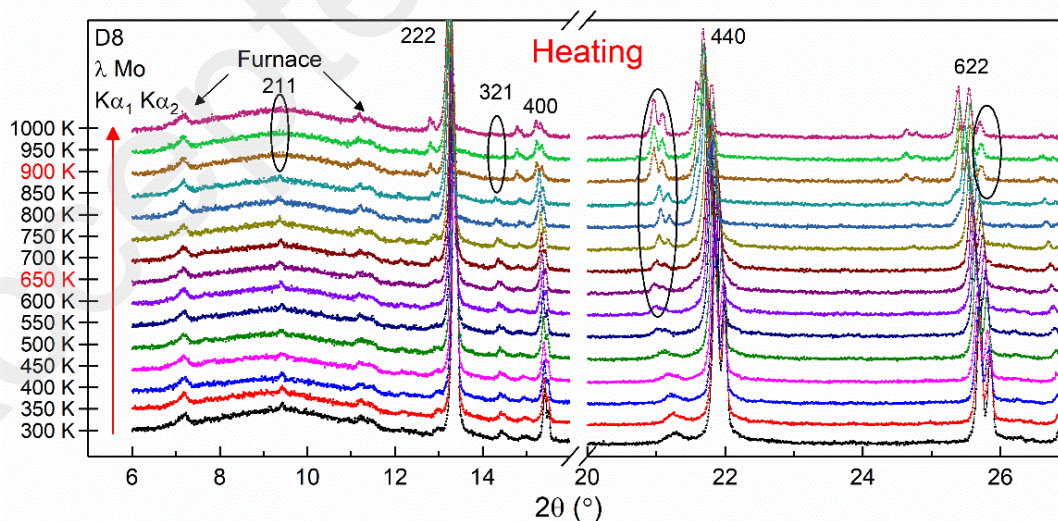


Figure 6. X-ray thermodiffractograms of germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ from RT to 1000 K ($\lambda = 0.71 \text{ \AA}$).

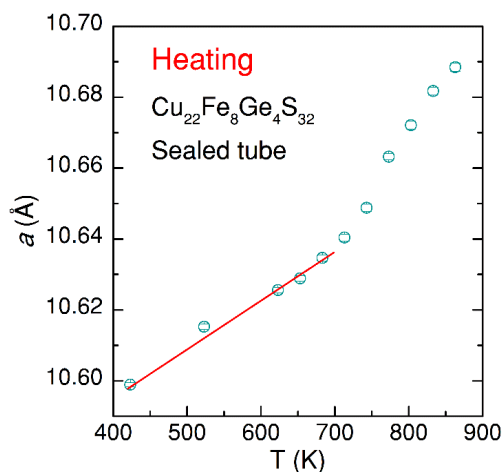


Figure 7. Thermal evolution of the unit cell parameter of germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ from Le Bail refinement of the XRD patterns recorded up to 860 K. Red line is just a guide for eyes.

Thus, from the interpretation of the combined thermal analyses we can conclude that (i) $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ germanite is chemically stable up to 650 K in inert atmosphere, (ii) above 650 K, a weak sulfur volatilization leads to the formation of Cu_3FeS_4 bornite and to a germanium enrichment of the germanite phase, and (iii) a cubic to tetragonal phase transformation occurs above 860 K, most probably related to the Ge enrichment of germanite.

CONCLUSION

This work demonstrates that *in situ* neutron powder diffraction is an efficient tool for the real-time determination of the natures and crystal structures of the intermediate phases occurring during sealed tube synthesis of complex sulfide materials. This study especially highlights the importance of the synthesis temperature and cooling method on the formation of high purity germanite $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ sample. The investigation of the phase stability of $\text{Cu}_{22}\text{Fe}_8\text{Ge}_4\text{S}_{32}$ shows that the formation of bornite is related to sulfur volatilization happening above 650 K. Also, it demonstrates that cubic germanite is highly sensitive to sulfur volatilization and transforms to a closely related tetragonal phase similar to renierite between 860 K and 900 K, due to germanium enrichment. This real-time investigation by *in situ* NPD can be easily applied to study and to optimize the synthesis conditions of other ternary and quaternary chalcogenides and opens the route to the discovery of new materials.

ASSOCIATED CONTENT

Supporting Information. Temperature profiles of *in situ* neutron powder diffraction experiments, detailed neutron and X-ray powder diffraction patterns, and specifications of the refinements with isotropic and anisotropic strain models. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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