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Nitromethane Pyrolysis in Shock Tubes and a Micro Flow Reactor With a Controlled Temperature Profile

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Abstract: Nitromethane has many applications, such as in racing, as a gasoline fuel additive, and as a monopropellant. Despite a large number of studies and the small size of the molecule, the combustion chemistry of nitromethane is still not well understood. To improve models, the pyrolysis of nitromethane (CH_3NO_2) was investigated experimentally in shock tubes and in a micro flow reactor with a controlled temperature profile (MFR), under dilute conditions. Several spectroscopic diagnostics were used in the shock tubes to follow the concentration time histories of CO, H_2O (both using IR laser absorption), and CH_3NO_2 (UV light absorption). A quadrupole mass spectrometer was used to measure CH_3NO_2 , NO_2 , CH_4 , C_2H_4 , and C_2H_2 at various temperatures with the MFR. These unique experimental results were compared to modern, detailed kinetics models from the literature, and no mechanism was able to reproduce these data over the wide range of conditions investigated. Predictions for the CO and H_2O levels were generally inaccurate, and the CH_4 , C_2H_4 , and C_2H_2 predictions were poor in most cases for the MFR data. Importantly, all models largely differ in their predictions. A numerical analysis was performed to identify ways to improve the next generation of nitromethane models. Results indicate that nitromethane decomposition needs to be improved below 1050 K, and that hydrocarbon-NO_x interactions still need to be further investigated.

Keywords: Nitromethane, pyrolysis, shock tubes, micro-flow reactor

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1. Introduction

Nitromethane (CH_3NO_2 , NM) can be used in several applications such as a racing fuel [1], a fuel additive for automotive applications [2,3], a reference fuel to understand the combustion of monopropellant [4,5] or a monopropellant itself [6]. Moreover, NM-filled cartridges can be used to generate a controlled deflagration to rapidly restore routes blocked by collapsed structures and rescue victims after earthquakes [7].

These applications, along with the fundamental interest in understanding the combustion chemistry of nitro compounds [8] (which can also help understanding better hydrocarbon/ NO_x interactions), explain why nitromethane combustion is receiving a lot of interest. One can mention studies in shock tubes [9-16], flames [1, 17-20], and plug flow and jet-stirred reactors [8,21,22]. Despite a large body of data for an apparently simple molecule, recent studies in flames or shock tubes showed that the combustion chemistry of nitromethane is still not well understood [12,17,20].

One good example can be found in the rapid thermal dissociation of CH_3NO_2 , which is one of the key features of its combustion [8,9,13,14,20,23-26]. Despite the importance of this dissociation, recent studies treat this step differently: Brequigny et al. [1] and Glarborg et al. [27] used only one decomposition reaction: $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3+\text{NO}_2(+\text{M})$ (R1) with the reaction rate from Glarborg et al. [28]. Mathieu et al. [9] used the reactions proposed by Annesley et al. [13] where the roaming isomerization of nitromethane was considered (resulting in two decomposition reactions: $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3+\text{NO}_2$ (R1) and $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3\text{O}+\text{NO}$ (R2)) whereas Weng et al. [22] revisited R1, used R2 from [13], and introduced another pathway: $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{HCNO}+\text{H}_2\text{O}$ (R3). Note that for the range of temperatures investigated herein, at 1 atm, the rate constant of the C–N bond cleavage R1 is higher than that of roaming isomerization (R2) and much larger than R3, which then can be ignored [22]. Finally, the most recent study of Shrestha et al. [8] considers only R1 and R2, but with reaction rates coming from Glarborg et al. [28] and Matsugi and Shiina [14], respectively.

More fundamental studies are required to improve NM models and, in the authors' opinion, it is now necessary to go beyond global kinetics data such as ignition delay times and laminar flame speeds. Accurate time histories and speciation of key species at various conditions are needed to further constrain the models, and a study of NM under pyrolysis conditions would also be beneficial to simplify the chemistry. The aim of

this study was to provide such fundamental data, assess modern detailed kinetics mechanisms, and identify ways to improve NM chemistry. To do so, pyrolytic mixtures of NM were investigated. CO and H₂O time histories were measured using laser absorption in shock tubes at Texas A&M University (TAMU); NM time histories were obtained by UV-light absorption at ICARE-CNRS Orléans; and various species were characterized by a quadrupole mass spectrometer (QMS) in a micro-flow reactor with a controlled temperature profile (MFR) at Tohoku University. First are presented the experimental conditions investigated and the experimental setups used. The results are then presented and compared with modern detailed kinetics mechanisms from the literature (Brequigny et al. [1], Shrestha et al. [8], Weng et al. [21], Glarborg et al. [27], and Mathieu et al. [29]). Finally, a model analysis was performed, and recommendations were made to improve models.

2. Experimental setups

The conditions investigated in the present work are summarized in Table 1, and wide ranges of temperatures, pressures, and NM concentration were covered. Note that the 1.5% NM condition was investigated with all apparatuses, between 1 and 2.2 atm.

Table 1: Summary of the experimental conditions investigated during this study.

| Reactor | Diagnostic | Mixture composition | Temp. (K) | Pressure (atm) |
|--------------------|--|--|-----------|----------------|
| Shock Tube (TAMU) | CO – laser absorption | 0.005 CH ₃ NO ₂ /0.2 He/0.795 Ar | 900-1380 | 1.55±0.25 |
| | | 0.015 CH ₃ NO ₂ /0.2 He/0.785 Ar | 900-1140 | 2.2±0.15 |
| | H ₂ O – laser absorption | 0.005 CH ₃ NO ₂ /0.995 Ar | 940-1250 | 1.05±0.15 |
| | | 0.015 CH ₃ NO ₂ /0.985 Ar | 1040-1240 | 2.25±0.25 |
| Shock Tube (ICARE) | CH ₃ NO ₂ – UV light absorption | 0.002 CH ₃ NO ₂ /0.998 Ar | 1040-1220 | 3.84-8.62 |
| | | 0.003 CH ₃ NO ₂ /0.997 Ar | 1060-1250 | 1.76-8.28 |
| | | 0.015 CH ₃ NO ₂ /0.985 Ar | 1060-1240 | 2.25±0.25 |
| MFR (Tohoku) | CH ₃ NO ₂ , CH ₄ , C ₂ H ₄ , C ₂ H ₂ , NO – QMS | 0.015 CH ₃ NO ₂ /0.985 Ar | 700-1300 | 1.0 |

Note that the CO formation was also measured for Helium-free mixtures. The addition of He was deemed necessary for vibrational relaxation, as detailed in Mathieu et al. [29]. However, another study from our group [30] did not show any appreciable benefits to adding He. In the present case, mixtures with and without He

presented similar time histories in terms of both timing and CO level. The data for the NM/He/Ar mixtures are presented here, but all the data for the NM/Ar mixtures are provided as Supplemental Material. Finally, the effect of impurities in the shock-tube results below was investigated using the work of Urzay and coworkers [31], showing that the impurities have virtually no effect on the profiles.

2.1 Laser absorption shock tubes

Laser diagnostics were employed in two different shock tubes at TAMU. These two facilities are relatively similar in that they are single-diaphragm, made of stainless steel, with a 7.62-cm i.d. driver section. Both have a larger driven section (15.24-cm i.d., 4.72-m long for the CO diagnostic; 16.2-cm i.d., 6.78-m long for H₂O) equipped with PCB P113A piezoelectric pressure transducers with a well-known spacing between them to measure the incident-wave velocities (determined at the endwall location with a curve-fit of the velocities measured between transducers). Post-reflected shock conditions were obtained with this extrapolated wave speed, in conjunction with the one-dimensional shock relations and the initial conditions in the test region. The uncertainty in the temperature determined behind reflected shock waves is within about 10-15 K and the time-accuracy of all the shock tubes used herein is considered to be about 1 μ s. The attenuation of monochromatic, cw laser light was used to determine the concentration of the target species using a Beer–Lambert relation: $I/I_0 = \exp(-k\nu P_{\text{abs}}L)$, where I_0 and I are the transmitted and incident laser light intensities, $k\nu$ is the spectral absorption coefficient, P_{abs} is the partial pressure of the absorbing species, and L is the path length. The linestrengths from the HITRAN 2004 database [32] were used to calculate $k\nu$, and I/I_0 was measured.

2.1.1. CO measurements

A quantum cascade laser was used to generate the light at 4566.17 nm to access the fundamental R(12) transition of the $v''=0$ CO band. The laser was centered on the R(12) transition using a removable CO/Ar absorption cell before each test. I and I_0 were monitored using two InSb detectors fitted with a bandpass filter centered at 4500 nm, with a full width of 500 nm, allowing for the broadband emission levels entering the

detectors to be $<0.3\%$ of the absorption signal. For more details on the CO diagnostic and the uncertainties in the measurements ($\pm 3.8\%$), see Mathieu et al. [29] and references herein.

2.1.2. H₂O measurements

Water formation was followed using the attenuation of light at 1388.140 nm coming from a Toptica Photonics tunable diode laser to access the $5_{5,1} \leftarrow 5_{5,0}$ transition at 1388.1389 nm within the v_1+v_3 fundamental band. The slight shift in wavelength is to account for the pressure shift of the line center. This diagnostic is described in more detail in Mathieu et al. [33], and the uncertainty on the water concentration is estimated to be below 6% [30].

2.2 UV Light absorption shock tube

The stainless-steel shock tube used at ICARE-CNRS Orléans is a double-diaphragm type, with a 1-m diver section (114.2 mm i.d.) and 3.75-m driven section (52.5 mm i.d.). The driven section is equipped with 4 shock detectors (CHIMIE METAL A25L05B) to determine the incident-wave velocity at the endwall location and related post-shock conditions, as described above. Measurements were situated 12 mm from the endwall through CaF₂ windows, with a light source (deuterium lamp HAMAMATSU C1518) and a monochromator (JOBIN-YVON HR250M, equipped with a HAMAMATSU R977 photomultiplier). The NM absorption spectrum was determined during the course of this study between 210.4 and 370.7 nm (spectrum available as supplemental material). Based on this spectrum, the absorption of NM was followed at 220 nm for most experiments, whereas some experiments performed in the highest pressure and temperature domains of the 1.5% NM mixture necessitated higher wavelengths (225-230 nm), to avoid saturation of the signal. Like for the laser-absorption measurements, the absorption intensity is proportional to the concentration of the absorbing species following the Beer-Lambert law. The absorption coefficient of NM in Ar was determined at each wavelength, for the entire range of temperatures investigated (visible in Sup. Mat.). Since the absorption of

products from NM decomposition cannot be fully excluded at 220 nm, only the first part of the profile (between 200 and 400 μ s for the highest and lowest temperatures, respectively) was considered.

2.3 Micro-flow reactor with a controlled temperature profile

The MFR is composed of a 2-mm-i.d. quartz tube heated by a hydrogen/air flat-flame burner providing a temperature ramp along the inner surface of the reactor channel. The burner can be moved along the reactor to vary the sampling position relative to the temperature profile. The maximum wall temperature (corresponding to the temperature of the inner surface of the reactor) was set at 1300 K during this study, and the measured wall temperature profile is shown in the figure presenting the results below. The uncertainty is < 0.3 mm for the axial distance, and the temperature difference between top and bottom walls inside the reactor is small (approx. 10 K), unaffacting the 1-D modeling. The NM/Ar mixture was supplied to the reactor at atmospheric pressure with an inlet mean flow velocity of 3 cm/s. A quartz micro-probe (0.1 mm i.d.; more details in [34]) was fused on the sidewall of the reactor channel, forming a T-shaped reactor. The micro-probe was used to sample the gases in the reactor channel, after which they are introduced into a QMS through a fused-silica capillary. The temperature of the sampling line was kept at 373 K to prevent water condensation. Chemical reactions in the sampling line were negligible thanks to the rapid pressure and temperature reductions of the sampled gas [34]. The following species were identified and quantified by QMS: NM, NO, CH₄, C₂H₂, and C₂H₄. Electro-ionization was used at an energy of 16 eV for CH₄ and NM, and 10 eV for C₂H₂ and C₂H₄, whereas photo-ionization was used for NO (10.8 eV) to minimize the fragmentation effects.

2.4 Modeling procedure

ANSYS Chemkin-Pro v19.0 was used to simulate the experiments. For the shock-tube results, the modeling was done using the closed homogeneous reactor module with U, V assumptions. For the MFR, the PREMIX code was used, where the flow field in the reactor channel was modeled as a one-dimensional, steady-state

reactive flow without a boundary layer. A heat convection term between the gas and the reactor wall was added to the gas-phase energy equation. The experimentally measured wall temperature profile was used during computations, and radical quenching was not considered. More details on the MFR and on the modeling procedure can be found in [35] and references therein. Note that radical quenching on the reactor surface is negligibly small in the MFR [36]. When missing, the transport data necessary to model the MFR results were assembled by using the AramcoMech 3.0 ($\text{H}_2/\text{O}_2/\text{Hydrocarbons}$) and the Glarborg models (NM and NO_x).

3. Experimental results and models comparison

3.1 CO Measurements

The CO formation from a 0.5% NM/diluents mixture, at around 1.5 atm, is visible in Fig. 1 for (a) low, (b) intermediate, and (c) high temperatures. As can be seen, the CO formation starts very rapidly and reaches a plateau value at a rate that increases with temperature. Note that the test time is not long enough for the low-temperature case to observe this plateau, and that the plateau value is different from the equilibrium value, not reached in the time frame considered. At low temperature, literature models are all notably under-reactive, with the Mathieu and Glarborg model being the closest to the data. At intermediate temperature, 1126 K, large differences between models can be observed. The Weng model is notably under-reactive and largely under-predicts the formation of CO by about 30%. To a lesser extent, a similar observation can be made with the Shrestha model (12%). The rapid, initial, growth in the CO signal is well-captured by the other models, although they are still slightly under-reactive, notably the Brequigny model.

Noticeable differences are also observed between models on the amount of CO reached during the test time. The Brequigny model predicts closely the amount of CO formed, whereas the Mathieu and Glarborg models over-estimate this value (by about 18 and 25%, respectively). For the high-temperature case, the Weng and Shrestha models are still under-predicting CO (10 and 25%, respectively), whereas the other models accurately predict the rise of the timing and slope of the initial CO formation. The Brequigny model slightly under-predicts the amount of CO at the inflexion point, but rapidly predicts the experimental profile with high

accuracy. Around the inflexion point, the models of Mathieu and, to a larger extent, Glarborg, over-predict the CO amount. These two models then eventually predict the same CO concentration after 1.2 ms, and the predictions and experimental data seem to be converging beyond the test time. Similar observations were made for the mixture with 1.5% NM, as can be seen in Supplementary Material.

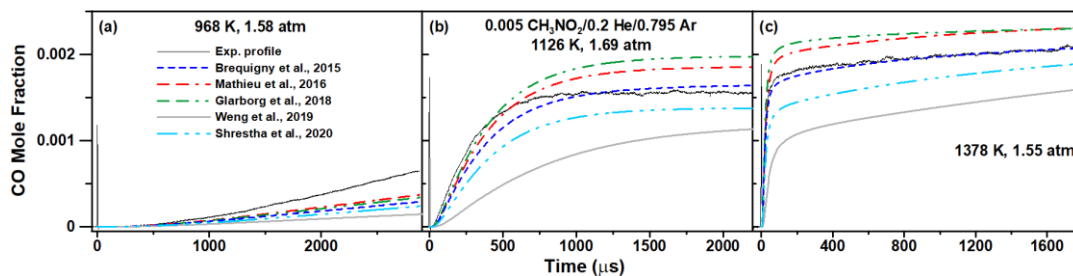


Figure 1: CO time histories and model comparisons for NM pyrolysis at (a) low, (b) intermediate, and (c) high temperature.

3.2 H₂O Measurements

The formation of water around atmospheric pressure for a 0.5% NM/Ar mixture can be seen in Fig. 2. Like for CO, the formation of H₂O is too slow in the time frame studied to reach its equilibrium for the lowest temperatures. At this condition, all models are under-reactive, with the Mathieu model being slightly better than the other ones. Concerning the Weng model, results always show an under-reactivity and a lesser formation of water, at all conditions (a 15% difference being observed at the plateau level of the high-temperature case (c)). As the temperature increases, the reactivity becomes similar for the three other models. The predicted level of H₂O is also rapidly exceeding the experimental observations, and an observation of the complete set of data shows that it occurs for temperatures above around 1075 K for the Glarborg and Brequigny models, and for temperatures above 1164 K for the Mathieu and Shrestha models. At high temperatures, Fig. 2c, the Mathieu and Shrestha models both over-estimate the amount of water by around 12-15%, whereas the Glarborg and Brequigny models over-estimate the water formation by around 20%. Like for CO, similar observations were made for the mixture with 1.5% NM (Supp. Mat.), although the amount of water predicted by the Brequigny

model is now visibly higher than for the Glarborg model for this higher NM concentration, with the Shrestha model being the most accurate in terms of H_2O concentration.

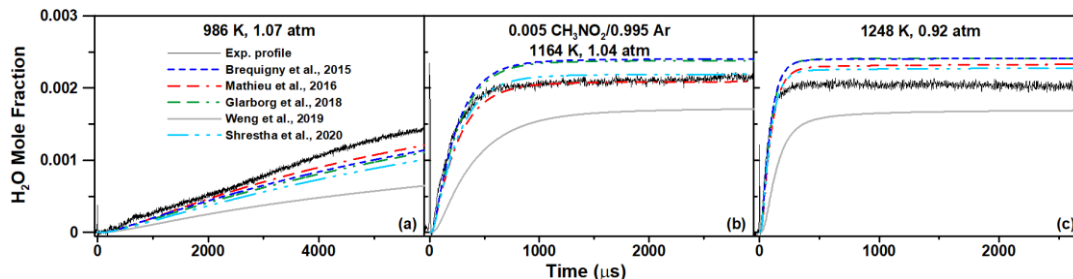


Figure 2: H_2O time histories and model comparisons for NM pyrolysis at (a) low, (b) intermediate, and (c) high temperature.

3.3 CH_3NO_2 Measurements

Some representative profiles for NM decomposition are visible for the most-dilute case (0.2% NM in Ar) at (a) low, (b) intermediate, and (c) high temperature in Fig. 3. As can be seen, less than 50% of the NM is decomposed within the time frame considered for the low-temperature case, whereas NM is fully decomposed by around 600 μs and 125 μs at intermediate and high temperatures, respectively. For these two last cases, predictions from all but Weng (under-reactive) models are accurate. However, while the Weng model is still under-reactive at low temperature, the four other mechanisms appear to be over-reactive. Similar results have been observed for the other conditions investigated, as visible in Supplementary Material.

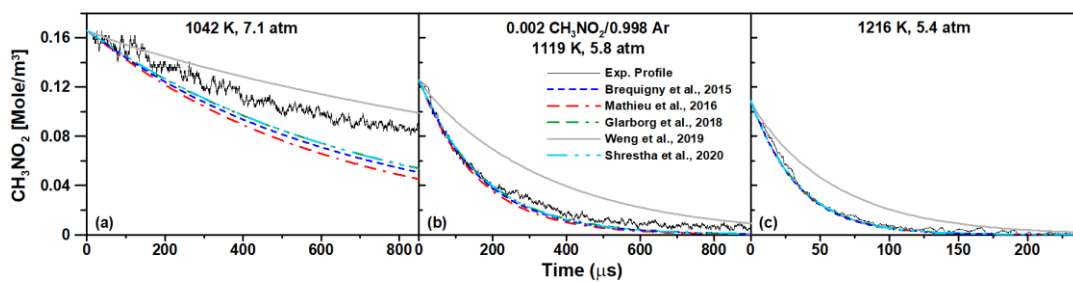


Figure 3: Comparison between experiments and models for CH_3NO_2 time histories during nitromethane pyrolysis at (a) low, (b) intermediate, and (c) high temperature.

3.4 Micro Flow Reactor measurements

The results for NM pyrolysis in a MFR are visible in Fig. 4. Concerning the NM profile (Fig. 4a), all models but, to a much lesser extent, the Brequigny et al. one predict a decomposition that is shifted towards lower temperatures (750 K/ $x = 5$ cm instead of 950 K/ $x = 5.4$ cm for the experiments, the Brequigny et al. model being almost within the experimental uncertainties). Parallel to the NM consumption, a large amount of NO is formed, up to 12300 ppm at around 1200 K (5.9 cm) corresponding to more than 80% of the initial NM concentration. As the temperature reaches its maximum, 1300 K (6.4 cm), about 1000 ppm of NO is consumed. The Brequigny model predicts the timing of the NO formation accurately, but the amount of NO is too large (14355 ppm). Moreover, the small consumption of NO is not predicted. A similar NO level and lack of NO consumption are observed by the other mechanisms, with the NO formation starting too early. However, while the model of Mathieu et al. also predicts early NO formation, the maximum amount of NO is within the experimental uncertainty, and this model is the only one capturing the small NO consumption past $x = 6$ cm (although this consumption is not large enough next to the experiment).

At the same time NM is consumed and NO is formed, a relatively large amount of CH_4 (up to 2360 ppm) is formed (Fig. 4b). As NO is consumed near $x = 6.5$ cm, it is visible on the experimental profiles that the CH_4 and C_2H_4 concentrations decrease in similar proportions, indicating that the two species are most likely interacting. Large differences in the predicted amount of CH_4 can be observed between the models, with the Brequigny model presenting a methane profile that is accurately reproduced until after $x = 6.5$ cm (the model actually presenting an opposite trend, with an increase in the CH_4 level). The Mathieu model also predicted the maximum amount of CH_4 within the experimental uncertainty, but with a too-early formation and with a CH_4 profile that increases, instead of the expected decrease, past 6 cm. All the other models predict the CH_4

formation too early, and in too large proportions, with the Glarborg model being the farthest (over-prediction by a factor of about 2).

Differences between the data and the models, and between the models, are larger for C_2H_4 (Fig. 4c), and C_2H_2 (Fig. 4d). For C_2H_4 , a noticeable decrease in the maximum amount is also observed past $x = 6.5$ cm, with most models predicting an opposite trend. This result and the opposite predictions for CH_4 past 6.5 cm could indicate that the CH_4/C_2H_4 and NO interactions need to be revisited in the models. All models predict the formation of ethylene starting too early, with the Glarborg et al. model being very close to the experimental data for the maximum C_2H_4 value registered. All other models over-predict the amount of ethylene, the farthest being the Brequigny model with an over-prediction by a factor larger than 5. On the other hand, acetylene is under-predicted by all models but the Mathieu et al. one. The late formation of C_2H_2 , past $x = 6$ cm, is not captured by the models (where the C_2H_2 formation starts at $x = 5.5$ cm or before), and so is the decrease in the C_2H_2 signal past $x = 7$ cm. The Brequigny model is within a factor 2 of the data, with the other models being worse, up to a factor of 11 for the mechanism from Glarborg and coworkers.

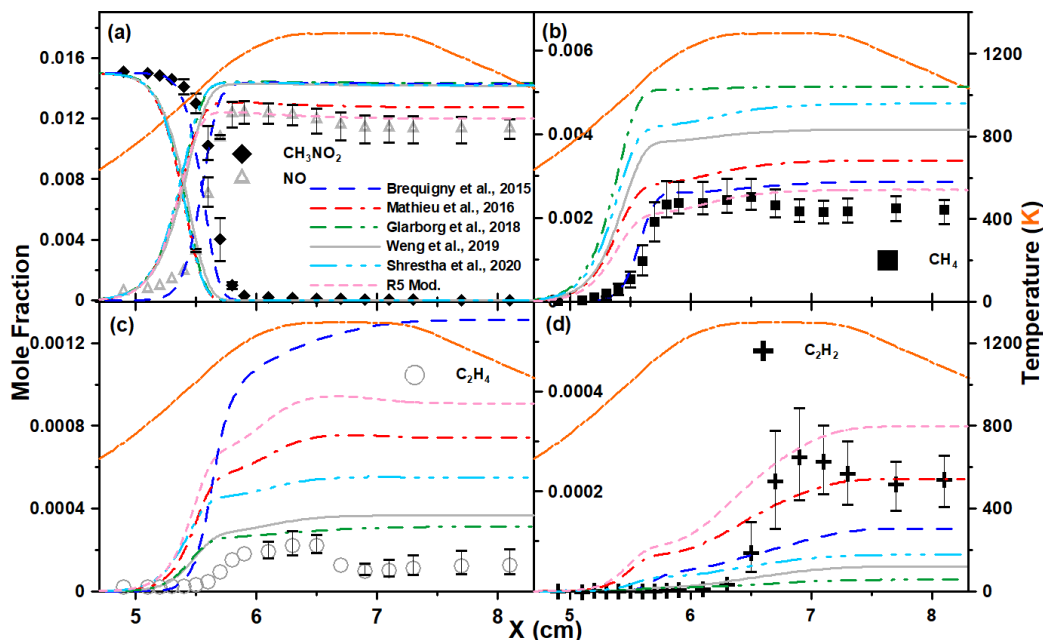


Figure 4: Comparison between experiment and models for (a) CH_3NO_2 and NO; (b) CH_4 ; (c) C_2H_4 ; and (d)

C_2H_2 during the pyrolysis of nitromethane in a MFR.

4. Discussion

The comparison between models and the data from this study demonstrated serious flaws in the pyrolysis chemistry of nitromethane. Notably, the decomposition of NM seems to be poorly predicted below 1050 K (MFR, shock tube), with most models being too reactive. On the other hand, issues at the NO_x-hydrocarbon interaction level were revealed: (i) the same models appear to be under-reactive for predicting CO and H₂O formation at these low temperatures; (ii) the predicted CO and H₂O levels were wrong in most cases; and (iii) the predictions for CH₄, C₂H₄, and C₂H₂ in MFR were poor. It can therefore be anticipated that the present data will be invaluable to help improving NM combustion chemistry and our current understanding of the interactions between NO_x and hydrocarbons.

To understand the reasons behind these poor predictions, a numerical analysis was conducted using reaction pathway, sensitivity, and rate-of-production analyses for the models considered herein. Concerning the NM decomposition, R1 is by far the most important reaction for all models, followed by R2 for the models that include this pathway. The particularity of the Weng model (Fig. 3) is due to the fact that it uses a different rate for R1 than the ones typically utilized, and that it has an additional unique reaction, with a relatively high sensitivity: $\text{CH}_3\text{NO}_2 + \text{CH}_3 \rightleftharpoons \text{CH}_3\text{NO} + \text{CH}_3\text{O}$ (R4) (their new NM decomposition pathway, R3, being unimportant under our conditions).

The MFR results are of particular interest, as the concentration levels of many species were poorly predicted and as a second reactive phase observed experimentally at $x = 6.5$ cm was not captured by any model. Using the Mathieu et al. model (as it offers decent predictions for NO concentration and is not too far from the data for CH₄), the numerical analysis shows the predominant role of the reaction $\text{HNO} + \text{CH}_3 \rightleftharpoons \text{NO} + \text{CH}_4$ (R5) in the formation of CH₄ and NO. Dividing the reaction rate of R5 (from Rasmussen et al. [37]) by a factor 3 brings the NO and CH₄ profiles in much better agreement with the data (“R5 mod.” in Fig. 4), but also increases the C₂H₄ and C₂H₂ levels. While C₂H₂ remains in relative agreement with the data, C₂H₄ is now significantly over-estimated. This change in the rate of R5 does not impact the shock-tube results herein. To further illustrate

the importance of this reaction, it is worth mentioning that using the reaction rate from Choi and Lin [38], used in the Glarborg and Shrestha models, largely increases the concentrations of CH₄ and NO, the latter being then identical to the prediction from the Glarborg model, the less accurate in this case (Fig. 4(c)). Note that reactions similar to R5 are not present in any model for molecules like C₂H₄ or C₂H₆, possibly leading to the poor predictions in the C2 species. Similarly, it is reasonable to think that the second reaction zone past 6.5 cm is not captured by the models due to missing interactions between hydrocarbon-type molecules and radicals with nitrogen-containing species.

Concerning the shock-tube results, the formation of CO will be used as an example to further illustrate the differences between models and to identify areas where the models need improvement. As seen in Fig. 1, a factor up to around 1.7 can be observed on the CO prediction at the plateau between the extremes, the Weng and Glarborg models. A rapid calculation shows that, when modeling the CO formation from a classical CH₄/O₂ mixture at $\phi = 2$, the models considered herein rapidly converge toward the same level of CO in the time frame investigated (Supp. Mat.), indicating that the differences observed between the NM models in Fig. 1 are not due to the base hydrocarbon chemistry.

An equilibrium calculation on CO was also performed at the conditions of Fig. 1b, and all models predict similar values for CO equilibrium (3272-3279 ppm), except the Weng et al. model which predicted a slightly lower value at 3210 ppm. This difference, certainly due to the thermodynamic properties used, is not large enough to explain the discrepancy between the Weng model and the others in terms of CO predictions. Interestingly, all models show a similar main reaction pathway for CO formation under our conditions: R1 (CH₃NO₂ ⇌ CH₃+NO₂) is always, by far, the main reaction behind NM consumption. R2 (CH₃NO₂ ⇌ CH₃O+NO), when present, or reactions between CH₃NO₂ and H or OH radicals to form CH₂NO₂ have a smaller contribution. In all models, the CH₃ coming from R1 will then react either with another CH₃ to form C₂H₆, or with NO₂ (R6) to form CH₃O and NO (like R2). Note that there are three different reaction rates used for R6, between the 5 models considered. The CH₃O from R6 (and R2) will then rapidly produce CO via the following sequence: CH₃O → CH₂O → HCO → CO. This sequence is the chief contributor

to CO in our experiments, and it is essentially controlled by reactions from the base hydrocarbon chemistry, except the transition from HCO to CO which occurs mainly via $\text{HCO} + \text{M} \rightleftharpoons \text{H} + \text{CO} + \text{M}$ (R7) and $\text{HCO} + \text{NO} \rightleftharpoons \text{HNO} + \text{CO}$ (R8).

For all models, R7 is, by far, the main reaction producing CO. In addition to having a similar main reaction pathway from CH_3O to CO, the rates of many of these reactions are similar between models. However, large discrepancies between the Glarborg and Weng models can be due to differences in the reaction rate selected for R7, and to two reactions contributing to CO formation present in the Glarborg model only: $\text{CH}_2\text{O} + \text{H} \rightleftharpoons \text{H} + \text{CO} + \text{H}_2$ (R9) and $\text{CH}_2\text{O} + \text{OH} \rightleftharpoons \text{H} + \text{CO} + \text{H}_2\text{O}$ (R10). Note that R10 also produces H_2O , which contributes to the over-predictions of water.

Concerning the C_2H_6 pathway, models also suggest a similar main reaction path: $\text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_4 \rightarrow \text{C}_2\text{H}_3 \rightarrow \text{C}_2\text{H}_2 \rightarrow \text{CH}_2\text{CO}$ after which CH_3CO and HCCO are formed in various proportions between models, to eventually form CO. This path is less important for CO formation, but contains many NO_x -hydrocarbon interactions. Judging by the level of predictions for C_2H_2 and C_2H_4 from the MFR results, as discussed earlier, these interactions are still not well understood and necessitate more work.

5. Conclusions

The pyrolysis of nitromethane was investigated experimentally in shock tubes by measuring CO, H_2O (IR laser absorption) and CH_3NO_2 (UV light absorption) time histories and in a MFR by measuring CH_3NO_2 , NO, CH_4 , C_2H_4 , and C_2H_2 (QMS). Results were compared with the most recent detailed kinetics models from the literature, and it was found that no model is capable of predicting these data accurately over the entire range of conditions investigated. While the predictions for NM consumption in shock tubes and the MFR were generally too fast in most cases below 1050 K, the CO and H_2O formation were too slow at these temperatures. In additions, there is no trend among these recent models, some predicting too much CO and/or H_2O , whereas some consistently under-predict these values. A numerical analysis with the models showed that the interactions

between hydrocarbons and NO_x need to be revisited and that some reaction pathways are very likely to be missing. One can conclude that the data from the present study will be extremely useful to develop a more accurate detailed kinetics model for nitromethane combustion.

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