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Single-template periodic mesoporous organosilica with organized bimodal mesoporosity

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Abstract

Periodic mesoporous organosilicas (PMOs) are obtained by the hydrolysis-condensation of organobridged triethoxysilane precursors in the presence of surfactants as structure-directing agents (SDAs). After removal of the SDAs, the resulting materials usually exhibit monomodal, well-organized mesoporosity, with the size of the pores being controlled by the SDA. However, despite the potential technological applications of such materials with well-organized bimodal porosity, to the best of our knowledge, PMOs exhibiting two distinct types of ordered mesoporosity have barely been described.

Herein, we describe a simple approach for modulating the dimensions of ordered monomodal 2D hexagonal and bimodal porosity in PMOs synthesized from 1,4-bis(triethoxysilyl)benzene (BTEB) and Pluronic P123 under acidic conditions, by varying the addition sequence of reactants. The approach employs a single SDA without degradation of the BTEB precursor. Reaction conditions leading to the formation of monomodal and bimodal porosity within the templated PMOs are identified. Our approach exploits the competition between the rates of (a) BTEB hydrolysis/condensation; and (b) diffusion, solubilization and partitioning of the unhydrolyzed and hydrolyzed precursor within the micelles. A mechanism describing the evolution of porosity within this system is proposed.

Keywords

Periodic Mesoporous Organosilica, Dual porosity, Bimodal porosity

1. Introduction

Highly porous materials based on silica have been prepared using a wide range of synthetic strategies and are now routinely exploited in many fields [1], including catalysis [2, 3], chromatography [4, 5], nuclear waste storage [6], insulation [7], optics [8], electronics [9], biocatalysis and biosensors [10-12]. The scope of such systems in applied materials science was profoundly impacted by the first reports of the successful preparation of periodic mesoporous silica (PMS) in the early 1990s [13], which exploited the use of surfactants as structure-directing agents (SDAs) to obtain pure mesoporous silica with well-organized and uniform pores. The framework structure and pore size/volume, together with the wall thickness, are modulated by careful selection of process parameters (including the nature of the SDA, use of swelling agents, temperature, solvent, etc), and the formation of several types of mesostructured silica, among which the cubic, 2D hexagonal and lamellar structures are most commonly obtained, has now been well-documented [14].

The utility of such organized mesoporous systems was further extended through the use of organosilane precursors (mono-silylated organics mixed with silica sources [15]) and by the first reports of the preparation of periodic mesoporous organosilica (PMO) analogues of PMS in 1999 [16-18]. A significant body of work has now emerged around control of the structure and properties of these latter materials for applications in such

diverse areas as optics, catalysis, nanomedicine, separation science, etc [19-30]. PMOs with high surface areas and uniform porosity are typically obtained via the hydrolysis-polycondensation of organosilanes with general stoichiometry $(R'O)_3Si-R-Si(OR')_3$ using appropriate SDAs. The organic functions (R) in the precursor are covalently bridged between two Si atoms, which ensures their regular distribution within the pore walls of the PMOs [31], while the range of potential bridging groups enables a correspondingly wide variety of functionalities to be introduced into the mesoporous hybrids [24, 32, 33].

A variety of SDAs have now been reported for the preparation of PMOs, which enable the pore size to be readily tuned from around 2 nm to more than 10 nm [34]. The SDAs employed include simple surfactants such as cetyltrimethylammonium bromide $((C_{16}H_{33})N(CH_3)_3Br)$, CTAB, Gemini molecules $(C_nH_{2n+1}N(CH_3)_2(CH_2)_sN(CH_3)_2C_nH_{2n+1})$ [35], triblock copolymers such as Pluronic P123 $((EO)_{20}(PO)_{70}(EO)_{20})$ [36] and polyion complex micelle species [29]. More sophisticated strategies involving the use of surfactant mixtures, co-solvents and co-surfactants as pore expanders [37], together with judicious control of pH, ionic strength and other key sol-gel process parameters [38-40] have also been reported. Well-organized porosity with controlled symmetry (e.g. 2D hexagonal, 3D hexagonal, cubic, etc), together with uniform pore sizes, can be readily obtained in such systems, with the range of applications critically dependent on both the functionality of the bridging species and the control of pore size and symmetry. A model describing the synergistic interactions between the SDA and precursor molecules during the evolution of templated porous structures has been proposed by Polarz and Antonietti [41]. In this so-called three-phase model, the hydrophobic domain of the surfactant and a depleted corona region at the

interface between the hydrophobic and hydrophilic regions control the size of the templated mesoporosity, while the hydrophilic chains typically interact with the evolving silicate or silsesquioxane framework and can contribute to the development of microporosity in the materials. The intrinsic hydrophobicity of the organic PMO precursors favors their solubility within the hydrophobic core and/or interface regions of the micelle, often leading to smaller pore sizes than would typically be observed for purely silica-based analogues [19].

A potential limitation of current approaches for producing such materials is that the pore size distributions obtained are typically monomodal, with few reports of synthesis strategies leading to bimodal porosity in PMOs having appeared [42-44] (in contrast to PMS-based systems [45-48]). However, many potential applications for such materials can be envisaged that would be greatly enhanced by the presence of well-organized bimodal porosity, including catalysis (control of the diffusion rates of reactants and products of different dimensions); chromatography (enhanced double-size exclusion); and controlled delivery of bioactive species (tailored release profiles of mixtures of pharmaceuticals); etc. Interestingly, bimodal porosity has also been obtained with the use of two different organosilanes. Indeed, such materials have been reported when (1,4-bis(triethoxysilyl)benzene (BTEB) and vinyltriethoxysilane are simultaneously added to the SDA solution [43]. A proposed mechanism suggests that the bimodality results from the vinyltriethoxysilane interaction with the corona of the SDA micelle, which increases the pore size. A similar result is obtained upon stepwise hydrolysis of two bridged organosilanes (1,2-bis(triethoxysilyl)ethene (BTSEENE) and BTEB), where mesoporous multipods were synthesized. It was demonstrated that the BTSEENE mesoporous nanorods grow onto the mesoporous spherical BTEB [49]. These complex

systems require the use of several organosilane precursors. When using only BTEB with P123 as surfactant, either 2D-hexagonal or cubic mesostructured PMOs are generally obtained [36, 42].

In this paper, we describe a simple approach for modulating the dimensions of ordered porosity in PMOs synthesized from BTEB and the symmetric triblock copolymer Pluronic P123 under acidic conditions, by varying the addition sequence of reactants. Reaction conditions leading to the formation of monomodal and bimodal porosity within the templated PMOs are identified. Our approach exploits the competition between the rates of (a) BTEB hydrolysis/condensation; and (b) diffusion, solubilization and partitioning of the unhydrolyzed and hydrolyzed BTEB sol-gel precursor within the three-phase regions of the self-assembling P123 SDA. A mechanism describing the evolution of porosity within this system is proposed.

2. Experimental

2.1 Chemicals

1,4-Bis(triethoxysilyl)benzene (BTEB) was synthesized using the previously described method [50, 51] and purified by distillation under vacuum. Pluronic P123 was obtained from Sigma Aldrich. Hydrochloric acid 35% (Rectapur), water (HPLC grade), ethanol (absolute grade) and acetone (Normapur) were obtained from VWR and used as received.

2.2 PMO synthesis

All syntheses were based on Inagaki's approach [36] and extracted according to Wang's method [52].

PMO-1: In a Schlenk vessel, Pluronic P123 (1.46 g, 218 μmol) was dissolved at 35 °C in water (53.5 mL) under strong agitation (1400 rpm). Concentrated HCl (310 μL , 3.5 mmol HCl) was added to the previous solution and the resulting mixture was then cooled to 0 °C. The BTEB (1.5 g, 3.7 mmol) was added dropwise to the cold solution. After 1 h of reaction, the reacting mixture was heated to 35 °C for 24 h and then to 100 °C for an additional 24 h. The white solid obtained was recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5 and then returned to the vessel in 53.5 mL of water at 100 °C for further aging and structural consolidation for three days. The resulting solid was then recovered by centrifugation, washed once with water, three times with acetone and dried under vacuum to afford 804 mg of white powder. The resulting material (550 mg) was extracted in an acidic solution of ethanol (4.3 mL of concentrated HCl in 143 mL of ethanol). The hybrid silica was finally recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5, then washed with acetone (three times) and dried under vacuum to obtain 485 mg of white powder. ^{13}C NMR (75 MHz, CP-MAS, ppm): 132.7 (aromatics), 58.6 ($\text{CH}_3\text{-CH}_2\text{-O}$), 17.0 ($\text{CH}_3\text{-CH}_2\text{-O}$). ^{29}Si NMR (60 MHz, CP-MAS, ppm): -62.0 (T^1), -71.0 (T^2), -79.4 (T^3). IR (cm^{-1}): 1036 ($\nu(\text{Si-O-Si})$), 914 ($\delta(\text{C}_{\text{sp}^2}\text{-H})$).

PMO-2: In a first Schlenk vessel, Pluronic P123 (1.46 g, 218 μmol) was dissolved at 35 °C in water (53.5 mL) under strong agitation (1400 rpm). Concentrated HCl (310 μL , 3.5 mmol HCl) was added to the previous solution and the resulting mixture was then cooled to 0 °C. In a second flame dried Schlenk flask, BTEB (1.5 g, 3.73 mmol) was added and cooled to 0 °C under magnetic stirring (750 rpm agitation). 20 mL of the

surfactant solution was then poured into the flask containing BTEB and the agitation speed was immediately increased up to 1400 rpm. 5 min later, the remaining surfactant solution was quickly added. After 1 h of reaction, the reacting mixture was heated up to 35 °C for 24 h and then to 100 °C for an additional 24 h. The white solid obtained was recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5 and then returned to the vessel in 53.5 mL of water at 100 °C for further ageing and structural consolidation for three days. The resulting solid was then recovered by centrifugation, washed once with water, three times with acetone and dried under vacuum to afford 839 mg of white powder. The resulting material (550 mg) was extracted in an acidic solution of ethanol (4.3 mL of concentrated HCl in 143 mL of ethanol). The hybrid silica was finally recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5, then washed with acetone (three times) and dried under vacuum to obtain 443 mg of white powder. ^{13}C NMR (75 MHz, CP-MAS, ppm): 132.6 (aromatics), 58.3 ($\text{CH}_3\text{-}\mathbf{\text{CH}_2}\text{-O}$), 16.4 ($\mathbf{\text{CH}_3}\text{-CH}_2\text{-O}$). ^{29}Si NMR (60 MHz, CP-MAS, ppm): -62.0 (T^1), -71.2 (T^2), -78.6 (T^3). IR (cm^{-1}): 1036 ($\nu(\text{Si-O-Si})$), 914 ($\delta(\text{C}_{\text{sp}^2}\text{-H})$)

PMO-3: In a first Schlenk vessel, Pluronic P123 (1.46 g, 218 μmol) was dissolved at 35°C in water (53.5 mL) under strong agitation (1400 rpm). Concentrated HCl (310 μL , 3.5 mmol HCl) was added to the previous solution and the resulting mixture was then cooled to 0 °C. In a second flame dried Schlenk flask, BTEB (1.5 g, 3.73 mmol) was added and cooled to 0 °C under magnetic stirring (750 rpm agitation). 18 mL of the surfactant solution was then poured into the flask containing BTEB and the agitation speed was immediately increased up to 1400 rpm. After 1 h, the temperature of the reaction mixture was increased to 35 °C for 24 h, and then to 100 °C for an additional

24 h. The white solid obtained was recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5 and then returned to the vessel in 53.5 mL of water at 100 °C for further ageing and structural consolidation for three days. The resulting solid was then recovered by centrifugation, washed once with water, three times in acetone and dried under vacuum to afford 764 mg of white powder. The resulting material (550 mg) was extracted in an acidic solution of ethanol (4.3 mL of concentrated HCl in 143 mL of ethanol). The hybrid silica was finally recovered by centrifugation, washed with water until the pH of the supernatant reached 6.5, then washed with acetone (three times) and dried under vacuum to obtain 451 mg of white powder. ^{13}C NMR (75 MHz, CP-MAS, ppm): 132.7 (aromatics), 58.3 ($\text{CH}_3\text{-CH}_2\text{-O}$), 16.2 ($\text{CH}_3\text{-CH}_2\text{-O}$). ^{29}Si NMR (60 MHz, CP-MAS, ppm): -62.7 (T^1), -71.5 (T^2), -79.5 (T^3). IR (cm^{-1}): 1039 ($\nu(\text{Si-O-Si})$), 914 ($\delta(\text{C}_{\text{sp}^2}\text{-H})$)

2.3 Characterization

^{29}Si and ^{13}C solid-state NMR spectra were obtained using cross-polarization and magic-angle spinning techniques (CP-MAS) on a Varian VNMR 300 MHz spectrometer. Thermogravimetric analyses (TGA) were performed on a TA Instruments Q50 apparatus with a heating rate of 10 °C.min⁻¹ under an air flow of 60 mL.min⁻¹. The material decomposition was studied from 25 °C up to 800 °C. Transmission electron microscopy (TEM) images were obtained on a JEOL 1400 P+ microscope operating at 120 kV with a LaB₆ source. Samples were prepared by microtome with a nominal thickness of around 70 nm. Fast-Fourier transform (FFT) analyses of selected regions of the TEM images were obtained using ImageJ software (Version 1.52a). Scanning Electron Microscopy (SEM) measurements were carried out using a Hitachi S4800 apparatus on platinum

metalized samples. Nitrogen physisorption isotherms were determined at 77.35 K using a Micromeritics ASAP 2020 machine. The activation of the materials was performed at 100 °C for 12 h under secondary vacuum. The specific surface area of the materials was determined using the BET model in the relative pressure range $0.05 < p/p^\circ < 0.25$, taking 0.162 nm^2 as the cross-sectional area for adsorbed nitrogen molecules. The α_s model was used for evaluating the presence of micropores and deriving the corresponding micropores volume, using Aerosil200 as reference.

The small- and wide-angle X-ray scattering (SWAXS) experiments were performed using a Guinier-Mering set-up coupled with a 2D image plate detector. The X-ray source was a molybdenum anode, delivering a high energy monochromatic beam ($\lambda = 0.71 \text{ \AA}$, $E = 17.4 \text{ keV}$) and providing structural information over scattering vectors q ranging from 0.02 to 2.5 \AA^{-1} . The region between the sample and the image plate was purged with flowing helium, to avoid air absorption. The data acquisition time was 3600 s and the glass capillaries used to hold the samples (Higenberg) had a thickness of 2 mm. The image azimuthal average was determined using FIT2D software from ESRF (France), and data corrections and radial averaging were performed *via* standard procedures.

3. Results and discussion

3.1 Structure and morphology of PMOs

The ^{13}C NMR spectra of all materials show similar profiles (Fig. S1, Supporting Information) and reflect the presence of the phenylene bridging group as well as small quantities of residual ethoxy groups. In the ^{29}Si NMR spectra, three signals corresponding to T^1 , T^2 and T^3 silsesquioxane moieties are observed in all cases with comparable relative intensities (Fig. S2, Supporting Information). The predominance of T^2 species indicates

that condensation is not complete and reflects the relatively slow kinetics of condensation under the acidic conditions used to synthesize the samples. Furthermore, no signals are visible below -100 ppm which indicates that no Si-C bond cleavage occurred. The ^{29}Si NMR spectrum of **PMO-2** is presented in Fig. 1.

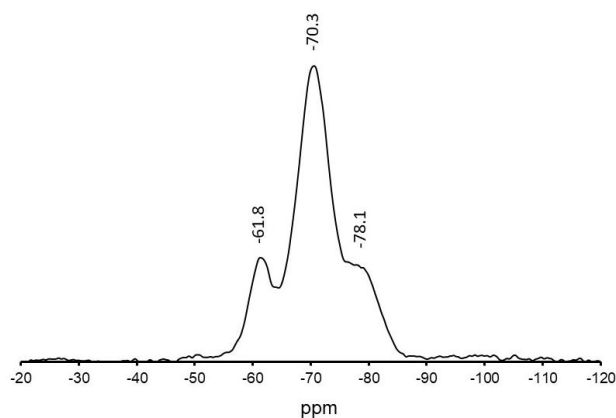


Fig. 1. ^{29}Si CP-MAS NMR spectrum of PMO-2

The FTIR spectra of **PMO-1**, **PMO-2** and **PMO-3** (Fig. S3, Supporting Information) exhibit the typical features expected for such hybrid materials [49]. Additional weak aliphatic $\nu(\text{C-H})$ modes are also evident from 2800 to 3000 cm^{-1} , which are attributed to the residual ethoxy species already identified by ^{13}C NMR (Fig. S1, Supporting Information).

Thermogravimetric analyses of the samples are illustrated in Fig. 2.

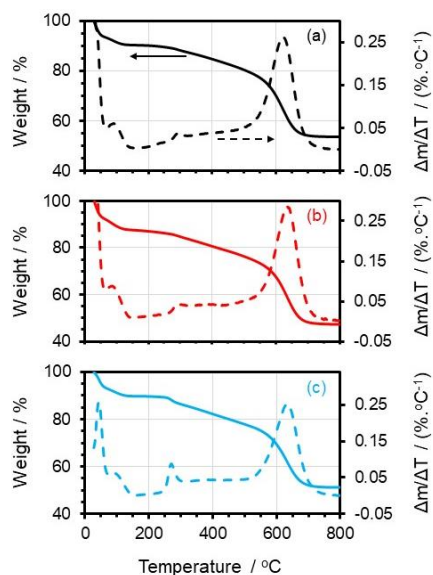


Fig. 2. Thermogravimetric analysis of PMO-1 (a); PMO-2 (b); and PMO-3 (c)

All samples exhibit similar features, which are consistent with their expected similar compositions. The plateau observed in the TGA data between 150 and 200 °C and weight loss (Δm) at temperatures below 150 °C are consistent with the initial removal of water (between 6.7 and 7.5 %), which occurs in two stages (based on the corresponding dTGA data). Below 75 °C, the water removal can be attributed to loss of residual “bulk” water retained within the pores after acid treatment, which is only weakly bound to the materials. Between 75 and 150 °C, a second stage is observed, which is attributed to the loss of water molecules more strongly bound to the structures. The subsequent degradation of the materials occurs above 250 °C, commencing with a small Δm at around 280 °C, a continuous and moderate Δm up to 600 °C and a significant Δm between 600 °C and 700 °C. The Δm at 280 °C (2 %) is tentatively attributed to the removal of residual organic species that were not released from the pores after acidic treatment. The continuous Δm (around 12 %) could correspond to the loss of geminal or isolated silanols [53]. In the same temperature range, it can be assumed that the degradation of the structure is also initiated through the oxidation of phenyl groups grafted onto the inorganic

framework, as previously reported for the degradation of hydrocarbons such as polymers or surfactants [54, 55]. Above 600 °C, the degraded organic species decompose. The residual quantities of SiO₂ above 700 °C are very similar, which indicates that the structuration of the three materials involves similar inorganic/organic molar ratios. This is consistent with the proportions of reagents used for the synthesis of the three materials.

The nitrogen sorption isotherms of the three materials are presented in Fig. 3.

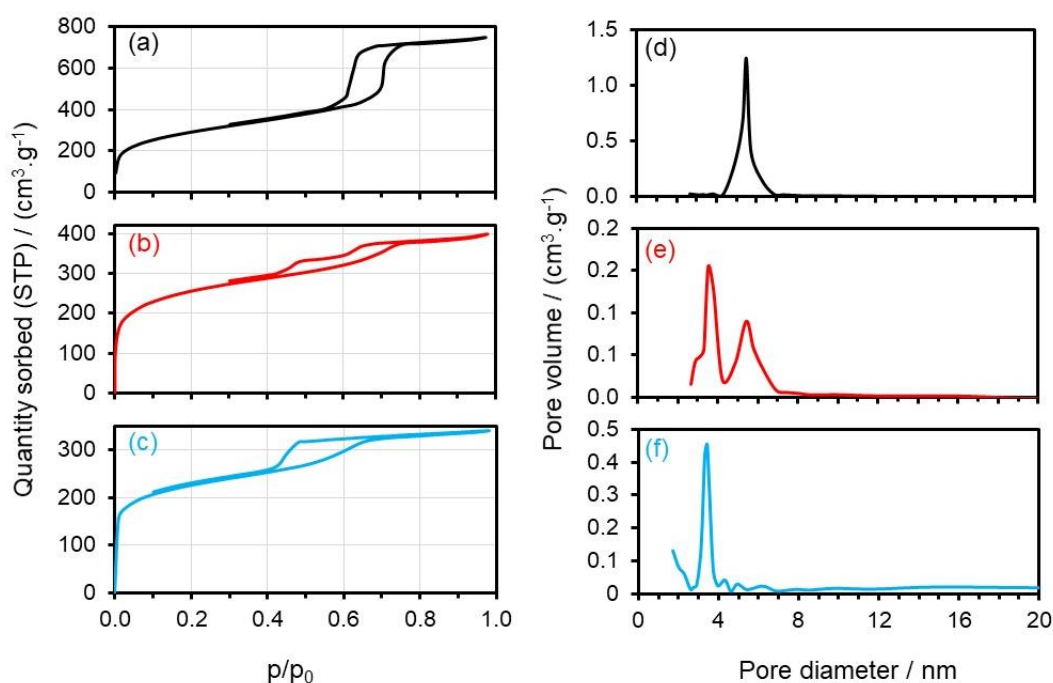


Fig. 3: Nitrogen physisorption isotherms of PMO-1 (a); PMO-2 (b); PMO-3 (c); and the corresponding BJH plots of PMO-1 (d); PMO-2 (e); and PMO-3 (f)

The curves obtained are typical type IV sorption isotherms, characterized by a strong uptake at intermediate relative pressure, followed by very flat plateaus. Additionally, the α_s plots, determined taking Aerosil 200 as a reference, revealed the absence of significant microporosity for these materials. Despite these common features, striking differences are evident. In the case of **PMO-1** (Fig. 3(a)), the hysteresis loop is characterized by two parallel branches, which is consistent with materials exhibiting well-organized mesoporosity. Interestingly, the high steepness of the uptake is an indication

that the mesoporosity is extremely homogeneous. This is confirmed by the BJH pore size distribution determined using the desorption branch of the sorption isotherm, starting from $p/p^\circ = 0.95$ for the derivation of the pore size distribution (Fig. 3(d)), where a very narrow peak centered at around 5.8 nm can be seen. In terms of pore volume, the capillary condensation process corresponds to about 50 % of the overall pore volume, which indicates that this material is essentially mesoporous, with very limited amorphous domains. This is confirmed by the high specific area obtained for this material, $1040 \text{ m}^2.\text{g}^{-1}$ (Table 1).

Table 1. Textural properties of PMO-1, PMO-2 and PMO-3

	PMO-1	PMO-2		PMO-3
		1 st population	2 nd population	
BET specific surface area (m^2/g)	1040		952	723
BJH pore size (nm)	5.8	5.8	3.9	3.9
SWAXS periodicity (nm)	11.3	11.5	16.1	11.5
Wall thickness (nm)	5.5	5.7	≥ 12	7.6-9.1
Pore volume (des., cm^3/g)	1.04		n/a	0.38

In the case of **PMO-2** (Fig. 3(b)), whereas the saturation plateau is still present, the uptake at intermediate relative pressure is less pronounced. Indeed, the saturation plateau is located at around $380 \text{ cm}^3.\text{g}^{-1}$ which is one half of the saturation plateau obtained for **PMO-1**. It can be deduced that this material is not as mesoporous as the material prepared using the pathway of **PMO-1**. On the other hand, the specific surface area is $952 \text{ m}^2.\text{g}^{-1}$, quite close to that obtained for **PMO-1**. This is consistent with the fact that the sorption isotherms for both samples are very similar at low relative pressure as shown in Fig. 4.

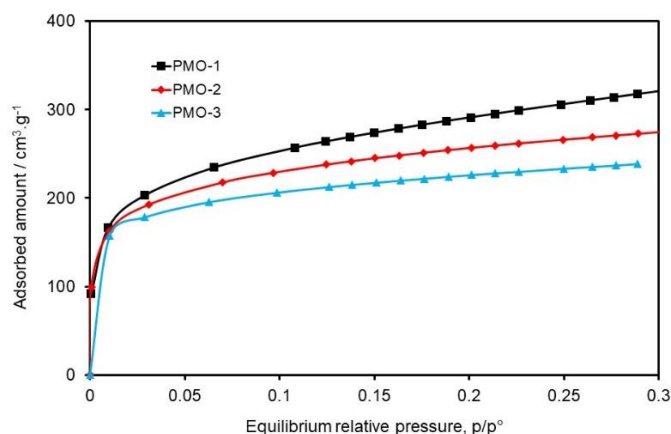


Fig. 4: Nitrogen adsorption isotherms at low relative pressure for PMO-1, PMO-2 and PMO-3.

It can be deduced that **PMO-1** has a larger population of mesopores, favoring larger amounts of condensed sorbate. In other words, the pore walls are likely thicker in the case of **PMO-2** (see discussion below and Table 1). Additionally, the uptake at intermediate relative pressure is not very steep, which is an indication that the material has mesopores of different sizes. More importantly, the shapes of the hysteresis loops are very different. In the case of **PMO-2**, the hysteresis branches are no longer parallel, and two desorption regimes can be distinguished. The shape of the hysteresis loop can be interpreted by determining the pore size distribution using the BJH method, obtained using the desorption branch. In Fig. 3(e), two peaks can be seen, which are attributed to two populations of mesopores with average diameters of 3.9 and 5.8 nm. It is noted that in this material, the latter pore diameter is identical to that found in **PMO-1**. Therefore, the data indicate that **PMO-2** exhibits bimodal mesoporosity, whereas **PMO-1** exhibits uniform, monomodal mesopores [36].

In the case of **PMO-3** (Fig. 3(c)), the saturation plateau is very close to that found with **PMO-2** (at around $350 \text{ cm}^3.\text{g}^{-1}$). However, a significant difference can be seen in the shape of the hysteresis loop. The branches are not parallel and there is only one

desorption step, as found with **PMO-1**. It can be deduced that **PMO-3** has only one population of mesopores, as confirmed by Fig. 3(f) where the BJH plot of the material is illustrated. Interestingly, the mean pore diameter is centered at 3.9 nm, which corresponds to one of the two pore sizes found for **PMO-2**. From the sorption analysis, it can be deduced that if **PMO-1** and **PMO-3** are mesoporous materials, with uniform, monomodal pore diameters related to the structuration of the materials, then **PMO-2** has two different populations of uniform mesopores. More importantly, these two populations of mesopores are found together in **PMO-2**, which exhibits a bimodal pore size distribution. This interpretation is consistent with the specific surface area obtained with **PMO-3**, $723 \text{ m}^2.\text{g}^{-1}$. Indeed, if **PMO-2** has a bimodal mesoporosity, its surface area should lie between those found for **PMO-1** and **PMO-3**, which is precisely the case.

The differences in the adsorption and desorption isotherm branches in the case of **PMO-2** (with the latter showing two distinct steps, in contrast to the single step seen for the former) can be attributed to percolation (bottle-neck) effects in such porous networks, as described by Rouquerol et al [56]. As two populations of mesopores are present in this material, the large pores will be emptied first, at a relative pressure which corresponds, through Kelvin's law, to the size of the smaller mesopores. Hence, a large decrease in the quantity of adsorbed gas can be seen in the first desorption stage. The smaller mesopores, in turn, will be emptied at lower relative pressure. In particular, the voids are emptied at relative pressures that satisfy the Kelvin law for void opening sizes, and hence, two steps are observed during desorption. This does not occur during the filling process, and hence only a single step is observed during adsorption. In addition, attempts to simulate the isotherms for **PMO-2** by using simple linear combinations of the isotherms for **PMO-1** and **PMO-3** did not identify any combinations that led to a good fit to the isotherms for

PMO-2. These data indicate that **PMO-2** is a unique material, and not a simple mixture of **PMO-1** and **PMO-3**.

The relatively thick pore walls observed in the case of **PMO-2** are also unusual, insofar as an increase in the thickness of the pore walls in such materials (Table 1) would normally be expected to result in a corresponding decrease in specific surface area and pore volume. However, in the present study, a comparison of the properties of **PMO-1** and **PMO-2** indicates that although the materials exhibit comparable surface areas, the wall thickness of **PMO-1** is much lower than that of **PMO-2**. This further highlights the fact that **PMO-2** with its bimodal porosity is a unique material, and a more complex system than **PMO-1** and **PMO-3**. The smaller mesopores contribute to the total specific surface area to a greater extent than they do to the total pore volume. In **PMO-2**, these effects interact to ensure that the increasing wall thickness and the presence of two populations of mesopores leads to a decrease in total pore volume without a corresponding significant decrease in overall surface area.

The SAXS spectrum of **PMO-1**, shown in Fig. 5 (bottom), exhibits four well defined pseudo-Bragg peaks in the low q region, at 6.4×10^{-2} , 11.2×10^{-2} , 12.9×10^{-2} and $19.6 \times 10^{-2} \text{ \AA}^{-1}$. This corresponds to q_0 , $\sqrt{3}q_0$, $\sqrt{4}q_0$, $\sqrt{9}q_0$, respectively, from a 2D hexagonal mesostructure ($p6mm$ symmetry) with a lattice parameter $a=11.3 \text{ nm}$ ($a = 4\pi/\sqrt{3}/q_0$).

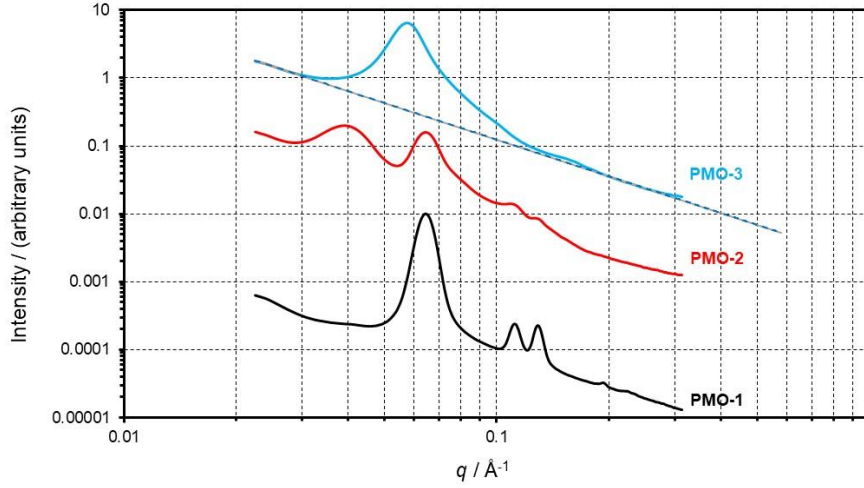


Fig. 5. SWAXS patterns for PMO-1 (bottom), PMO-2 (middle) and PMO-3 (top)

The TEM image confirms the 2D hexagonal phase (Fig. 6(a), see below), with a measured lattice parameter of 11.5 ± 0.1 nm which is also in agreement with the SAXS data. The apparent wall thickness, calculated from the lattice parameter a and the pore diameter determined *via* nitrogen adsorption (5.8 ± 0.5 nm), is 5.5 ± 0.6 nm.

The SAXS spectrum of **PMO-3** (Fig. 5, top) displays a relatively strong, broad peak at $5.7 \times 10^{-2} \text{ \AA}^{-1}$ (q_0 , corresponding to a characteristic distance of 11 nm), together with a weaker peak at around 0.17 \AA^{-1} ($3q_0$). The shape of the scattering profile suggests the presence of an additional broad peak (or peaks) at around 0.1 \AA^{-1} . On the basis of these data alone, it is not possible to unambiguously identify the symmetry of the uniform pore network in **PMO-3**, since the results are consistent with either hexagonal or cubic symmetry. However, FFT analysis of the TEM data from **PMO-3** (Fig. 6, see below) suggests that the symmetry is hexagonal, with a corresponding lattice parameter $a = 12.7$ nm ($a = 4\pi/\sqrt{3}/q_0$). Hence, the apparent wall thickness, based on the pore diameter deduced from the BJH analysis (3.9 nm), is ~ 8.8 nm (assuming hexagonal symmetry), which is significantly thicker than for **PMO-1**.

In contrast, the SAXS spectrum of **PMO-2** (Fig. 5, middle) is more complex. Three peaks at 6.4×10^{-2} , 11×10^{-2} and $13 \times 10^{-2} \text{ \AA}^{-1}$ are assigned to Bragg reflections from the (100), (110) and (200) planes under $P6mm$ symmetry, clearly demonstrating the 2D hexagonal structure with the same lattice parameter as **PMO-1** (11.3 nm). This is further confirmed by TEM measurement (see below), from which a lattice parameter of $11.5 \pm 0.5 \text{ nm}$ is obtained (Fig. 6(c)). However, an additional low- q signal observed at $3.9 \times 10^{-2} \text{ \AA}^{-1}$ (q_0 , corresponding to a characteristic distance of 16 nm) suggests the presence of a second uniform pore network, as also observed via TEM (see below). This bimodal material has a first wall thickness, calculated from the lattice parameter of the $P6mm$ space group and the pore diameter determined *via* nitrogen adsorption ($5.8 \pm 0.5 \text{ nm}$), at $5.5 \pm 0.5 \text{ nm}$ associated with the 2D hexagonal phase, together with a larger wall thickness (pore diameter=3.9 nm) exceeding 12 nm associated with the second phase. In this system, SWAXS, nitrogen physisorption and TEM measurements (see below) are all consistent with the presence of two uniform mesoporous domains.

It should also be noted that in the WAXS region (high q values), characteristic distances at $d=7.6 \text{ \AA}$ and 3.8 \AA (q value of 0.8 and 1.7 \AA^{-1} , respectively) are observed for all samples (exemplified by the data for **PMO-1** in Fig. S4), consistent with a crystalline-like framework with a molecular-scale periodicity. The first peak corresponds to the lamellar repetition of the aryl moieties in organosilane fragments, whereas the second is assigned to the distance between the phenylene bridges, which are ordered *via* π - π interactions [22, 57].

TEM images obtained from the PMOs are illustrated in Fig. 6 and Fig. S5, together with FFT data obtained from selected regions of the images. The micrograph of **PMO-1**

and corresponding FFT (Fig. 6(a)) exhibit monomodal, organized porosity with hexagonal symmetry, consistent with the SWAXS and N₂ sorption data presented above.

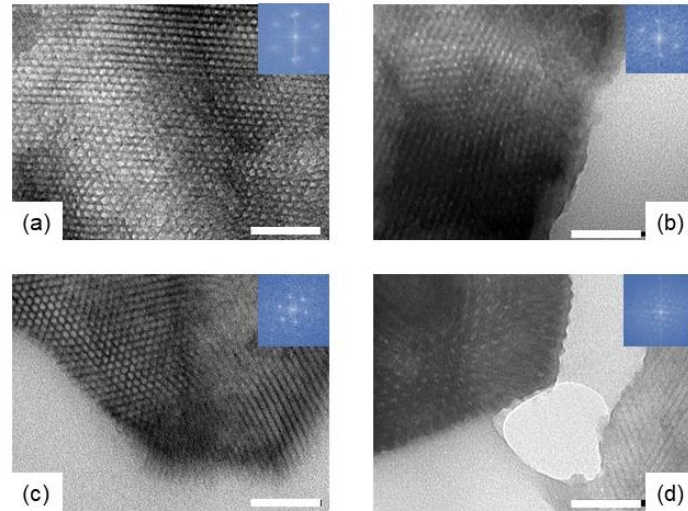


Fig. 6. TEM micrographs of PMO-1 (a); PMO-3 (b) and PMO-2 (c,d). (Size bar: 100 nm). Inset: Fourier transforms of selected regions, highlighting pore symmetry

PMO-3 (Fig. 6(b)) also exhibits uniform, monomodal porosity with thicker pore walls than in the case of **PMO-1**, although the structure appears somewhat less ordered. The corresponding FFT data (inset to Fig. 6(c)) are also consistent with hexagonal symmetry. In contrast, the structuring of **PMO-2** (Fig. 6(c), (d)) is substantially different, with two different uniform pore structures observed in different regions of the samples. Fig. 6(c) shows pores with uniform, organized hexagonal symmetry similar to those observed in Fig. 6(a) for **PMO-1**, with a comparable wall thickness (Table 1). On the contrary, Fig. 6(d) reveals a second population of uniform pores which appear less organized than those in Fig. 6(c), although the corresponding FFT data (inset to Fig. 6(d)) are consistent with hexagonal symmetry [14]. The wall thickness of these latter pores is substantially higher than that observed for either **PMO-1** or **PMO-3** (Figure 6(d) and Table 1). In addition, the TEM data presented in Fig. S5 indicates that both populations

of mesopores identified by N₂, SWAXS and TEM in the case of **PMO-2** are co-located in the same particles, further highlighting the unique nature of **PMO-2** and confirming that it is not a simple mixture of **PMO-1** and **PMO-3**. As indicated above, these data are entirely consistent with the SWAXS and N₂ physisorption data.

SEM micrographs of the PMOs are included in Fig. 7. The structuring observed on these length scales for **PMO-1** (Fig. 7(a)) and **PMO-2** (Fig. 7(b)), which both exhibit organized 2D hexagonal mesoporosity, is similar to that reported by Goto and Inagaki [36], and is consistent with a pseudo-crystalline material. In contrast **PMO-3** appears substantially less ordered (Fig. 7(c)), as expected from the earlier discussion.

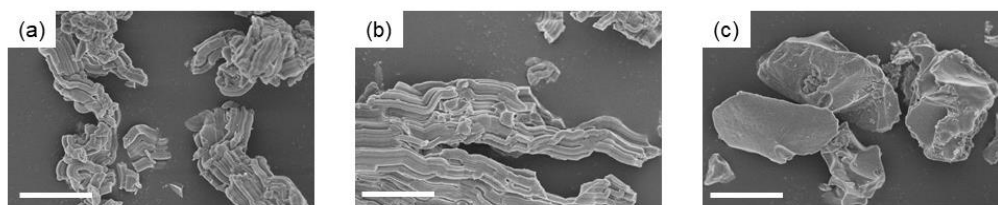


Fig. 7. SEM micrographs of PMO-1 (a); PMO-2 (b) and PMO-3 ©. (Size bar: 5 μm)

3.2 Mechanism

A notable feature of the different morphologies observed in this study is that they are obtained simply by varying the addition sequence of reactants at 0 °C during the initial 5 min of reaction at that temperature, without changing precursors and using a single surfactant. A consistent processing regime is then applied to all three systems, which involves aging of the initial structures at 35 °C for 24 h, followed by an additional aging step at 100 °C for a further 24 h. As previously discussed, the ²⁹Si-NMR data (Fig. 1) indicates that the extent of condensation is essentially identical in all three samples, with predominantly T² silicon species observed. Similarly, the thermogravimetric analyses

(Fig. 2) suggests that the sample stoichiometries are similar. Hydrolysis is expected to be relatively fast under these conditions, with only minimal condensation expected (estimated pH ~ 1.2) [58]. These observations suggest that the basis for the observed morphologies is largely established at 0 °C by the complex interplay between the key reaction parameters. These include hydrolysis of BTEB, π - π stacking, solubilization and partitioning of the hydrolyzed and unhydrolyzed BTEB within the various domains [41] of the self-assembled (2D hexagonal) P123, H-bonding between EO (and PO) groups in the P123 chains, etc. Subsequently, increasing the reaction temperature to 35 °C would lead to a significant increase in the rate of the relatively slow condensation step. As a consequence, the formation of the hybrid silica networks from the kinetically stable structures initially established at 0 °C will be accelerated. The network will be further consolidated by aging for an additional 24 h at 100 °C. The key differences in the processes occurring during the initial mixing step at 0 °C are explored in more detail below and summarized schematically in Fig. 8.

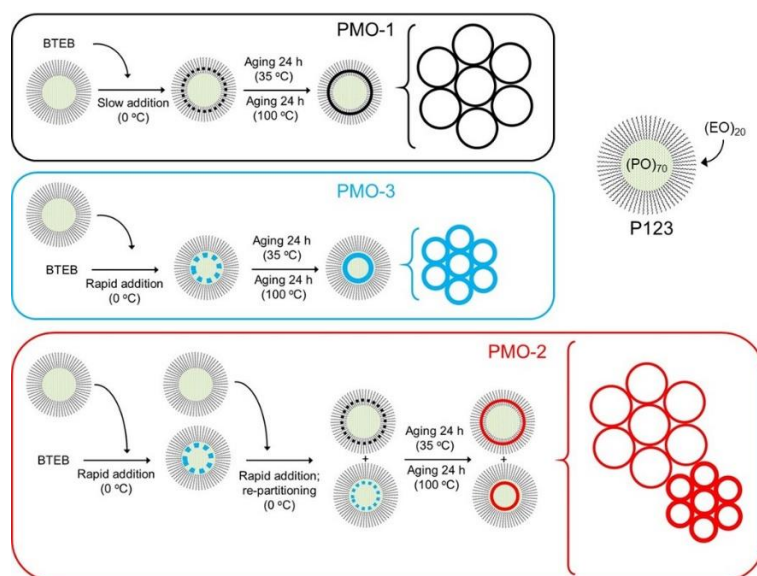


Fig. 8. Proposed mechanism for formation of uniform mesoporosity in PMO-1, PMO-2 and PMO-3

3.2.1 PMO-1

PMO-1 was prepared under conditions reported in earlier studies, *i.e.* by dropwise addition of the BTEB precursor to an acidic P123 solution [36]. The BTEB:H₂O:HCl:P123 mole ratio is 1:800:0.95:0.07 after complete mixing of the reagents, and hence water is in large excess at all times during the reaction. A simple calculation suggests that the pH is around 1.2, so hydrolysis is expected to be relatively fast and condensation relatively slow [58], as indicated above. It would be anticipated that the BTEB molecules would be rapidly hydrolyzed under the prevailing conditions, generating hydroxylated intermediates that would be less hydrophobic than the initial BTEB species. Hence, they would be expected to accumulate outside of the hydrophobic core of the self-assembled P123 structures and within the corona region, due to the competition between H-bonding of the silanol groups attached to the hydrophilic (EO)₂₀ P123 groups and solubilization of the phenylene species within the hydrophobic ((PO)₇₀) core. Cerveau *et al* [59] reported the formation of hexasilanol species formed in biphasic systems by the controlled hydrolysis of BTEB. These are stabilized by H-bonding and π - π stacking to form well-organized supramolecular species without condensation between the hexasilanol species. Such organized supramolecular species, together with smaller hydrolyzed BTEB species, would further self-assemble by π - π stacking, EO-silsesquioxane H-bonding interactions and subsequent condensation to form the PMO network. This network formation would thus be expected to proceed mainly within the corona region and in close proximity to the hydrophobic (PO)₇₀ core, thus establishing the size of the pores in the evolving PMO network and during the subsequent aging/consolidation steps at 35 and 100 °C. The SAXS pattern for **PMO-1** (Fig. 5) indicates that the pores within the sample have 2D hexagonal symmetry, consistent with

results previously reported for such materials in several studies [19, 36]. In addition, the N₂ sorption data indicates that there is essentially no microporosity in the sample, suggesting that the various aging and consolidation steps (including the three-day treatment at 100 °C) generates well-ordered, uniform mesoporosity with few micropores.

3.2.2 *PMO-3*

The mixing sequence used to prepare **PMO-3** is the reverse of that for **PMO-1**, insofar as the P123 solution is rapidly added to neat BTEB in the former case, albeit with only 1/3 of the quantity of surfactant being used. Under these conditions, the BTEB:P123 ratio is increased three-fold compared to **PMO-1** and the entire quantity of BTEB is essentially unhydrolyzed when the reactants are fully mixed. It would be expected that the more hydrophobic BTEB species would penetrate further into the hydrophobic (PO)₇₀ core of the micelles prior to hydrolysis commencing. Subsequent hydrolysis of the BTEB species should then lead to repartitioning of the BTEB moieties between the corona (H-bonding with EO groups) and hydrophobic core (where both solubilization of the phenyl groups and H-bonding with the PO groups can occur). Under such conditions, the formation of well-organized supramolecular structures such as those discussed above [59] would be less favorable, leading to poorer long-range ordering in the evolving PMO. In addition, the higher concentration of BTEB would be more likely to promote some condensation and formation of small oligomers, which would also inhibit the formation of well-organized supramolecular structures. Due to the lower quantity of P123 used to prepare **PMO-3** compared to **PMO-1** (three-fold reduction), it is also evident that more silane species would be associated with each P123 micelle in the case of **PMO-3**. During subsequent condensation at 35 °C to form the PMO network, the deeper initial penetration

of BTEB precursors into the core of the self-assembled P123 aggregates would be expected to lead to a single population of smaller pores in the aged material compared to **PMO-1**. In addition, the larger quantity of silane precursor associated with each P123 aggregate would be expected to lead to thicker walls in the final product, as is observed.

It is evident from a comparison of the SAXS patterns of **PMO-1** and **PMO-3** that the uniform pore network in **PMO-3** is not as ordered as that in **PMO-1** (with only a single, broad peak being clearly observed, Fig. 5). This suggests that the hydrolyzed BTEB precursor is not as well organized in **PMO-3** prior to condensation commencing, compared to **PMO-1**, as discussed above.

3.2.3 *PMO-2*

The addition sequence in the case of **PMO-2** is similar to that of **PMO-3**. In a first step, an aliquot of the acidic P123 solution is added to neat BTEB, and hence the initial reactions occurring would be equivalent to those occurring in the case of **PMO-3**, with penetration of BTEB into the micelle core, H-bonding of partially hydrolyzed BTEB with both EO and PO groups, etc. However, after 5 min, an additional aliquot of acidic P123 solution was added in the case of **PMO-2**, to yield a final BTEB:P123 ratio comparable to that in the case of **PMO-1**. Since the extent of condensation under these conditions is limited [58], it would be anticipated that some of the now-hydrolyzed BTEB species associated with the initially added P123 aggregates in the **PMO-2** system would re-partition into the unoccupied P123 species incorporated during addition of the second aliquot. However, unlike the conditions prevailing immediately after the addition of the first P123-solution aliquot (where the BTEB is unhydrolyzed and relatively hydrophobic), significant hydrolysis of BTEB would have occurred by the time of the

addition of the second aliquot. Hence the conditions prevailing during re-partitioning resemble those for **PMO-1**, where the BTEB species interacting with unoccupied P123 micelles would be at least partly hydrolyzed, with the possible formation of small oligomers. Hence, the BTEB species would be expected to partition into the corona region of P123 micelle added in the second aliquot and to further self-assemble. During subsequent consolidation of the structure via condensation at 35 °C, a bimodal pore-size distribution would thus be established, with the pore sizes resembling those observed for **PMO-1** and **PMO-3**. This expectation is consistent with the SAXS and N₂ sorption data presented in Figs. 5 and 3, respectively, which reveals the presence of a uniform, well-ordered population of pores with 2D hexagonal symmetry (as observed for **PMO-1**), together with a second population of uniform pores with diameters equivalent to those found in **PMO-3**.

4. Conclusion

PMOs with uniform, organized 2D hexagonal monomodal and bimodal porosity have been prepared under acidic conditions from BTEB, using a simple approach involving a single SDA (Pluronic P123). In contrast to earlier reports, the method described herein does not require the use of co-surfactants or multiple SDAs and does not involve decomposition of the bridging organic species to generate the bimodal mesopore network. Instead, the addition sequence of the reactants and SDA is controlled to exploit the competition between key process parameters: hydrolysis, π - π stacking, solubilization/partitioning of the hydrolyzed and unhydrolyzed BTEB within the various domains of the self-assembled (2D hexagonal) P123, H-bonding between EO and PO

groups in the P123 chains, etc, at 0 °C. Subsequent aging at 35 to 100 °C consolidates the silsesquioxane networks to generate the observed products.

The approach outlined opens up new possibilities for producing functional materials with uniform, well-organized bimodal pore networks that retain the functionality of the bridging organic species within the walls of the materials. Many potential applications for such materials can be envisaged that would be greatly enhanced, including catalysis (controlled diffusion rates of reactants and products of different dimensions); chromatography (enhanced double-size exclusion); and controlled delivery of bioactive species (tailored release profiles of mixtures of pharmaceuticals); etc.

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