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ARTICLE

Design of stable mixed-metal MIL-101(Fe/Cr) materials with enhanced catalytic activity for the Prins reaction

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This work highlights the benefit of designing mixed-metal (Cr/Fe) MOFs for enhanced chemical stability and catalytic activity. A robust and stable mixed-metal MIL-101(Cr/Fe) was prepared through a HF-free direct hydrothermal route with Fe³⁺ content up to 21 wt%. The incorporation of Fe³⁺ cations in the crystal structure was confirmed by ⁵⁷Fe Mössbauer spectrometry. The catalytic performance of the mixed metal MIL-101(Cr/Fe) was evaluated in the Prins reaction. MIL-101(Cr/Fe) exhibited a higher catalytic activity compared to MIL-101(Cr), a higher chemical stability compared to MIL-101(Fe) and a higher catalytic activity for bulky substrates compared to MIL-100(Fe). In situ infra-red spectroscopy study suggests that the incorporation of Fe³⁺ ions in MIL-101 structure leads to an increase in Lewis acid sites. It was thus concluded that the predominant role of Cr³⁺ ions was to maintain the crystal structure, while Fe³⁺ ions enhanced the catalytic activity.

Introduction

Metal-organic frameworks (MOFs) have been successfully applied as catalysts for a large variety of liquid-phase reactions.¹⁻⁵ Of particular importance are the Lewis acid-catalyzed processes due to the wide reaction range and their general applicability for the production of fine chemicals, pharmaceuticals and agrochemicals.⁶ Some examples have shown the superiority as catalysts of MOFs *versus* zeolites to promote Lewis acid-catalyzed reactions in the liquid phase.^{7,8} Among the most commonly used MOFs as heterogeneous catalysts, MIL-101(Cr) combines a series of desirable properties, including: i) an exceptional chemical and a suitable thermal stability, ii) a very high porosity with wide pentagonal (1.2 nm) and hexagonal (1.4 nm) windows to access to the

mesoporous cavities (2.9 and 3.4 nm in diameter), iii) a large surface area (3000-4200 m² g⁻¹), and iv) a tunable composition through organic ligand substitution and metal node modification.⁹ Importantly, after the removal of water molecules coordinated to the metal sites, the metal node cluster (Cr³⁺-trimer) presents inherent Lewis acidity.⁷ Thus, MIL-101(Cr) has been widely reported as heterogeneous catalyst for those reactions that require Lewis acid sites^{7,10} or redox¹¹⁻¹³ centers.

Heterogeneous catalysis is continuously searching for more active and selective catalysts. In the particular case of MOFs as solid Lewis acids, this target can be achieved, by at least, two alternative ways, either by introducing electron-withdrawing substituents on the organic linker or by selecting metal ions with higher acid strength. In this context, it has been reported that the catalytic activity as Lewis acid of MIL-101(Cr) can be increased up to three orders of magnitude by introducing electron-withdrawing groups such as nitro groups in the organic ligand.¹⁰ However, a complementary strategy would be to partially replace Cr³⁺ by other appropriate transition metal. Regarding the possibility to enhance Lewis acidity by proper selection of the transition metal, an obvious choice is the replacement of Cr³⁺ by Fe³⁺ in MIL-101. It has been previously established that for some reactions MIL-101(Fe) is more efficient as catalyst than MIL-101(Cr),¹⁴ the direct comparison between the catalytic activity of Cr³⁺ and Fe³⁺ is in most cases not possible due to the different nature of the accompanying counter anion, particles sizes and other parameters. In addition, MOFs bearing Lewis acid sites such as MIL-100 or MIL-101 suffer from a partial poisoning of the most active metal sites by impurities (ligand, anion...)¹⁵ Thus depending on

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the activation conditions, the same material can exhibit a significantly different density of Lewis sites of different strength,¹⁶ making thus comparison between various samples cumbersome.

In this context, the lower chemical stability of MIL-101(Fe), easily converted upon exposure to water or polar organic solvents to the less porous MIL-88B(Fe),¹⁷ has hampered the determination of the real catalytic activity of Fe³⁺ ions in the MIL-101 structure in comparison with the Cr³⁺ analogue.

Previously some of us have reported the direct preparation of mixed Cr/Fe MIL-53 materials. It was shown that the introduction of Cr³⁺ ions into the Fe-MOF was accompanied by a strong enhancement of the chemical stability.¹⁸ Later on this strategy was further extended to prepare highly porous Cr-MOFs through an easy solvothermal post-synthetic treatment starting from pre-formed Fe-MOFs.¹⁹ However, the synthesis of bi-metallic kinetic phases, such as MIL-101, remains challenging as it requires a careful tuning of the kinetic reactivity of the cations.

Preparation of heterometallic MOFs is a well-established strategy to optimize the catalytic activity of these materials.²⁰⁻²³ For instance, doping MIL-100(Fe) with Ni species was found a convenient way to both increase the activity of the MOF catalyst, while preserving the good chemical stability of the bare Fe-MOF.

Herein, the preparation of a robust and stable mixed-metal MIL-101(Cr/Fe) structure in various Fe³⁺ proportions up to 21 wt% through a direct hydrothermal synthesis avoiding the use of toxic HF. ⁵⁷Fe Mössbauer spectrometry supports the successful incorporation of Fe³⁺ cations within the crystal structure of the mixed-metal MIL-101(Cr/Fe) solids. Catalytic data for the Prins reaction indicate that the mixed-metal MIL-101(Cr/Fe) materials exhibit both an increased activity in comparison with the pure MIL-101(Cr) material, as well as a strongly enhanced chemical stability compared with the single-metal MIL-101(Fe) solid. The mesoporosity of the mixed-metal MIL-101(Cr/Fe) was found to make this material more adequate as catalyst to promote the Prins addition of bulky substrate compared to MIL-100(Fe), most likely due to a favored substrate diffusion. In addition, a good relationship between the Fe³⁺ proportion in the MIL-101(Cr/Fe) catalysts and the observed catalytic activity has been found. Stability data indicate that MIL-101(Cr/Fe) solid acts as a true heterogeneous catalyst that can be reused maintaining its initial crystallinity without undergoing metal leaching. This work illustrates the successful design of heterometallic MOFs exhibiting a superior catalytic activity by having one metal with a predominant role in maintaining the crystal structure (Cr³⁺ in the present case) and a second one enhancing the catalytic activity (Fe³⁺ in MIL-101).

Results and discussion

Catalyst preparation

The synthesis of MIL-101(Cr/Fe) by direct hydrothermal synthesis has been already reported.²⁴ However, the

previously reported synthetic procedure used hydrofluoric acid (HF) as acidic modulator/mineralizing agent. HF is a highly hazardous liquid (highly toxic and corrosive, as well as contact poison) and, therefore, not suitable for green and sustainable syntheses that are targeted nowadays, particularly for large scale preparation. A hydrothermal synthesis without HF was recently reported, but pure MIL-101(Cr/Fe) phase was only obtained for low Fe content (10 wt%).²⁵ In the present case, the target was to introduce much larger Fe proportions and about 20 wt% of the total Cr(NO₃)₃·9H₂O was replaced by Fe(NO₃)₃·9H₂O. Consequently, the reported synthesis was optimized to avoid the use of HF, while achieving higher Fe content.

For the synthesis of pure and well crystallized MIL-101(Cr/Fe) solids, one of the most difficult challenge to overcome is the competition between the crystallization of two other polymorphs (i.e. same chemical composition, but different structure) typically obtained when dealing with M³⁺ cations and terephthalic acid: namely MIL-101,⁹ MIL-88²⁶ and the thermodynamically preferred phase MIL-53.^{27, 28} The use of mixed-metals with different reactivity is also known to modify the crystallization process and influence polymorph formation, not mentioning the risk to form metal oxides.²⁹ Therefore, attempts to obtain mixed-metal MIL-101(Cr/Fe) using iron nitrate as Fe³⁺ source and similar synthetic conditions (220 °C, 8 h) as those used for the pure MIL-101(Cr),³⁰ resulted only in mixtures of MIL-101/MIL-88B (Figure SI-1), possibly due to the higher reactivity of iron cations that drives the phase formation towards MIL-88B. The addition of non-toxic acidic modulators, such as acetic acid (HOAc) was also explored, but it resulted as well in MIL-101/MIL-88B mixtures (Figure SI-1). A possible way to address this difference in reactivity between Fe³⁺ and Cr³⁺ is the use of a less reactive iron source such as metal iron. Fe⁰ has already successfully been used for the preparation of the mixed metal MIL-53(Cr/Fe).²³ As seen in Figure SI-2, the use of Fe⁰ drastically limited the formation of MIL-88B phase, compared to the use of iron nitrate, but did not fully prevented its formation. The use of HOAc did not improve the phase purity and resulted in mixture of MIL-101/MIL-88B phases (Figure SI-2). It was, thus, concluded that this acid modulator was not adequate for the formation of a pure MIL-101(Cr/Fe) phase.

Basic additives, such as NaOH, KOH and tetramethylammonium hydroxide (TMAOH), have also been studied for the synthesis of MIL-101 materials. Their use has been reported to enhance the solubility of the terephthalic acid and to favor the nucleation process, as the pH conditions are shifted around pH 6, thus promoting the synthesis of MIL-101. MIL-101(Cr) synthesized in the presence of low concentrations of TMAOH showed a good crystallinity, high specific surface area and the formation of pure phases.^{31, 32} It was hence of interest to investigate the use of this basic additive for the formation of the mixed metal MIL-101(Cr/Fe). To minimize the formation of the thermodynamic phase MIL-53(Cr), the reaction temperature was decreased, while the reaction time increased and various TMAOH concentrations were tested (0.0125-0.09 M). It was found that the purity of

MIL-101 (Cr/Fe) varied linearly with TMAOH concentrations (Figure SI-3). Concentrations above 0.06 M promoted mostly the synthesis of MIL-101 phase, whereas lower concentrations resulted in traces of the MIL-53 phase. When TMAOH concentration was increased above 0.07 M, the crystallinity of the product significantly decreased until an amorphous solid was obtained at 0.09 M of TMAOH. These results are in agreement with previous studies that showed the effect of the TMAOH concentration on the crystallinity of MIL-101.³²

Further optimization was achieved by further decrease of the reaction temperature minimizing oxide formation. Two pure MIL-101(Cr/Fe) materials with 4:1 and 5.9:1 Cr:Fe ratios were finally successfully obtained by using TMAOH as basic additive at 0.06 M at 150 °C. Both MIL-101(Cr/Fe) materials showed characteristic Bragg peaks of MIL-101, without any additional peaks, which confirmed the formation of a pure phase (Figure 1).

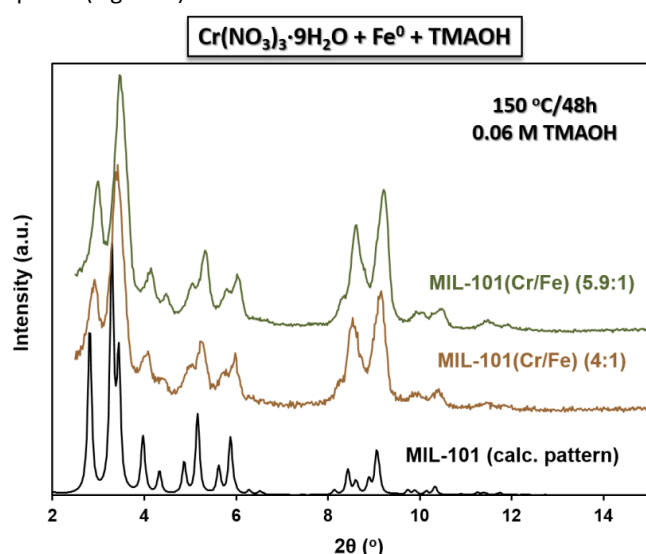


Figure 1: Normalized PXRD patterns ($\lambda = 1.54 \text{ \AA}$) of MIL-101(Cr/Fe) (4:1) and MIL-101(Cr/Fe) (5.9:1), compared with the calculated pattern of MIL-101(Cr).

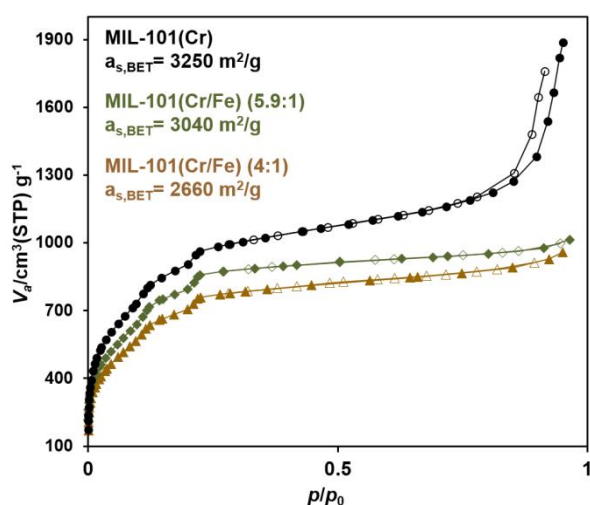


Figure 2: N_2 sorption isotherms of MIL-101 (Cr) (black cycles), MIL-101(Cr/Fe) (4:1) (brown triangles) and MIL-101(Cr/Fe) (5.9:1) (green diamonds) at 77 K ($P_0=1 \text{ atm}$).

The Cr and Fe atomic content was analyzed by EDX (average of 12 random points) and confirmed by ICP-OES. The composition was found to be $81 \pm 2 \%$ Cr and $19 \pm 2 \%$ Fe for MIL-101(Cr/Fe) (4:1, 21 wt%) and $85 \pm 1 \%$ Cr and $15 \pm 1 \%$ Fe for MIL-101(Cr/Fe) (5.9:1, 14 wt%). The low standard deviation as well as the identical values obtained by EDX and ICP-OES, suggest a homogeneous dispersion of iron in the sample. The chemical formulae of the MOFs deduced from the EDX are $\text{Cr}_{2.4}\text{Fe}_{0.6}\text{OH}(\text{H}_2\text{O})_2\text{O}[(\text{O}_2\text{C})\text{-C}_6\text{H}_4\text{-(CO}_2)]_3$, $\text{Cr}_{2.57}\text{Fe}_{0.43}\text{OH}(\text{H}_2\text{O})_2\text{O}[(\text{O}_2\text{C})\text{-C}_6\text{H}_4\text{-(CO}_2)]_3$ for MIL-101(Cr/Fe) (4:1) and MIL-101(Cr/Fe) (5.9:1), respectively

SEM analysis on the two mixed-metal MIL-101(Cr/Fe) solids revealed the formation of particles, without a well-defined morphology. Nonetheless, both MOFs showed uniform particles of around 100 nm (Figure SI-4). The TGA curves of the two mixed metal MOFs were compared with that of MIL-101(Cr) (Figure SI-5). The thermal degradation of the structures was observed around 300 °C under oxygen, which is similar to the pure MIL-101(Cr),⁹ indicating that the incorporation of Fe^{3+} cations, does not affect the thermal stability of the framework. The percentage of Cr_2O_3 residue obtained in the case of MIL-101(Cr) (31 %) was found to be slightly lower than then global percentage of the oxide residues ($\text{Cr}_2\text{O}_3 + \text{Fe}_2\text{O}_3$) formed in the case of the mixed metal MOFs: 35 % for MIL-101(Cr/Fe) (4:1) and 32 % for MIL-101(Cr/Fe) (5.9:1). This may indicate that a small part of the metal precursors used for the synthesis of the MOFs could have formed oxides that could not be removed during the activation of the samples. The N_2 sorption measurements of the mixed-metal MOFs exhibit type I isotherms, with secondary uptakes at $p/p_0 \sim 0.1$ and $p/p_0 \sim 0.2$, which are characteristic of the two microporous windows (pentagonal and hexagonal) of the two mesoporous cages (Figure 2). The apparent BET surface areas were calculated to be 2700 ± 100 and $3040 \pm 60 \text{ m}^2/\text{g}$ for MIL-101(Cr/Fe) (4:1) and MIL-101(Cr/Fe) (5.9:1), respectively. Both mixed metal MOFs exhibited slightly lower surface area than that for the pure MIL-101(Cr) ($3250 \text{ m}^2/\text{g}$), which could be related to the presence of small percentages of metal oxides, in agreement with the TGA measurements.

^{57}Fe Mössbauer spectrometry was used to gain information about the oxidation and the spin state of Fe, along with its electronic environment, in order to confirm that the Fe of the mixed MOFs was indeed incorporated in the lattice and to exclude the possibility of Fe being mainly under the form of Fe oxides or hydroxides. The transmission Mössbauer spectra of MIL-101(Cr/Fe) (4:1) recorded at 300 K and 77 K, consist of quadrupolar doublets with broadened and overlapped lines, indicative of different environments of Fe atoms (Figure 3, SI-6 and SI-7). They were first recorded at 4 mm/s to check that the samples did not contain any High Spin (HS) Fe(II) species and then at 2 mm/s. Figure 3 illustrates only the spectra recorded at 2 mm/s. Different fitting models could be applied with two components or three components (Figure SI-6), resulting in the same mean values of isomer shift (0.33 and 0.48 mm/s at 300 K and 77 K, respectively) and quadrupolar splitting (0.68 and 0.74 mm/s at 300 K and 77 K, respectively). These values are

consistent with the presence of HS Fe^{3+} in an octahedral environment, suggesting that they may be located in the octahedral units of the inorganic building blocks.³³⁻³⁵ This is further supported by the similarity between the Mössbauer spectra of the mixed MIL-101(Cr/Fe) (4:1) and the reported spectrum of MIL-100(Fe) (Figure SI-7 and Figure 3).³⁵ The isomer shift values of MIL-100(Fe) are significantly higher (0.42 and 0.54 mm/s at 300 K and 77 K, respectively). This difference could result from the amount of fluorine ions contained in the structure, as the presence of F ions in the Fe environment results in higher isomer shifts than Fe surrounded by O.²³

It should be noted that in the reported Mössbauer spectrum of MIL-100(Fe), the synthesis was performed in presence of KF and it has been previously demonstrated that the resulting structure has one fluorine atom per Fe trimer.³⁵ However, in the case of MIL-101(Cr/Fe), the synthesis was performed without fluorine, and the corresponding coordination position in the structure is occupied by hydroxyl ions.¹⁸ In the present mixed metal MIL-101(Cr/Fe), the mean values of isomer shift are rather typical of Fe^{3+} surrounded by oxygen or hydroxyl groups, in fair agreement with the absence of KF during the synthesis of MIL-101(Cr/Fe). When the content of Fe incorporated in the synthesis was decreased (ratio Cr/Fe 5.9:1), no significant differences were observed in the Mössbauer spectrum (Figure 3), except for a slight increase of the quadrupolar doublet asymmetry. The presence of water molecules in the framework actually influences the Mössbauer spectra and may be responsible for this difference. It is important to emphasize that the lack of resolution of the hyperfine structure does not allow excluding the presence of traces of iron³⁺ oxides in MIL-101(Cr/Fe) samples, similarly to the case of MIL-100(Fe).

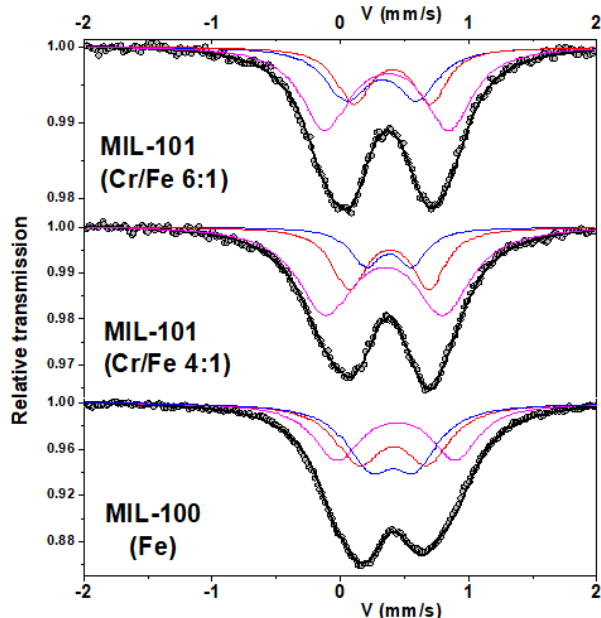


Figure 3: Mössbauer transmission spectra of MIL-101(Cr/Fe) (5.9:1), MIL-101(Cr/Fe) (4:1) and MIL-100(Fe), recorded at 77 K.

Our strategy therefore allows the preparation of MIL-101(Cr/Fe) solid with a Fe^{3+} loading up to 21 wt% using a HF-free route. It should be commented that the mesoporous rigid MIL-101 phase exhibits higher BET surface areas and porosity ($3250 \text{ m}^2 \text{ g}^{-1}$ and $1.2 \text{ cm}^3 \text{ g}^{-1}$) compared to that of the flexible microporous MIL-88B ($<500 \text{ mg}^2 \text{ g}^{-1}$; $0.6 \text{ cm}^3 \text{ g}^{-1}$), whose pores undergo a strong contraction upon drying. Therefore, MIL-101 is better suited as heterogeneous catalyst in comparison to MIL-88B, which shows important diffusion limitations.

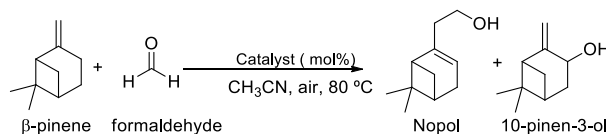
Catalytic activity

As commented in the introduction, the main purpose of this study was to show that it is possible to overcome the intrinsic structural instability of the single metal MIL-101(Fe) solid by exploiting the structural stability of MIL-101(Cr), while introducing in nodal positions some Fe^{3+} sites in adequate proportion to preserve structural stability due to the prevalence of Cr^{3+} , but still sufficiently high to achieve optimal catalytic activity.

To prove the concept of one metal maintaining the MOF structure and other acting as catalytic center, the MIL-101(Cr/Fe) (4:1) sample having about 21 wt% Fe^{3+} was initially selected and its catalytic activity studied for the benchmark Lewis acid catalyzed addition of β -pinene to formaldehyde to afford [2-(7,7-dimethyl-4-bicyclo[3.1.1]hept-3-enyl)-ethanol], commonly known as Nopol (Scheme 1). Nopol is an important commodity with woody camphoraceous scent used in perfumery for the large scale production of fragrances.^{8,36} It is also used in the preparation of pesticides in the agrochemical industry.^{8,36} This Prins reaction has been widely studied in the presence of a large range of homogenous (i.e. HCl, SnCl_2 , ZnCl_2 etc.) and heterogeneous catalysts including MOFs, observing that Nopol formation is accompanied by the presence of some byproducts, 10-pinen-3-ol being the most common one.⁷

Particularly, formation of Nopol has been reported already using MIL-100(Fe) or MIL-100(Cr) as heterogeneous catalysts, observing that these materials efficiently promote the addition reaction, without undergoing metal leaching and exhibiting reusability in consecutive batch reactions.⁸

In the present study a preliminary blank control experiment in the absence of catalyst shows that β -pinene conversion does not take place under the present reaction conditions. In contrast, when MIL-101(Cr) was present in the reaction, the addition reaction occurs resulting in the formation of Nopol as main product together with 10-pinen-3-ol, reaching a conversion of 26 % at 30 h (Figure 4). The use of MIL-101(Fe) or MIL-88B(Fe) as iron-based catalysts resulted in a slight increase of the catalytic activity ($\sim 36\%$ β -pinene conversion at 30 h) respect to the use of MIL-101(Cr).



Scheme 1. Prins reaction of β -pinene with formaldehyde and the chemical structure of the two products observed in the present reaction.

PXRD patterns of the used MOFs revealed that MIL-101(Fe) is, however, unstable under the present reaction conditions and is largely transformed into the MIL-88B(Fe) structure (Figure SI-8). Importantly, the use of the novel MIL-101(Cr/Fe) (4:1) solid resulted in a much more robust material, exhibiting a significantly higher catalytic activity (85 % at 30 h) than MIL101-(Cr), MIL-101(Fe) or MIL-88B(Fe) and even than the homogeneous Fe^{3+} acetate catalyst (Figure 4).

To put the catalytic activity of MIL-101(Cr/Fe) (4:1) into context, MIL-100(Fe) with all the metal nodes constituted by iron was used as reference. MIL-100(Fe) exhibited only slightly lower activity than MIL-101(Cr/Fe) (4:1), whose Fe^{3+} content is about 5 times lower. Thus, it seems that the Fe^{3+} sites of MIL-101(Cr/Fe) (4:1) are intrinsically more active than the ones of MIL-100(Fe). The lower intrinsic activity of Fe^{3+} in MIL-100(Fe) can be understood considering the smaller dimensions of windows (0.55 and 0.86 nm) and cavities (2.4 and 2.9 nm) in MIL-100 respect to the MIL-101 structure imposing diffusion limitations to the reaction. The higher activity of MIL-101(Cr/Fe) (4:1) respect to MIL-101(Cr) can also be understood considering the benefits of combining the highly porous and stable MIL-101(Cr) structure, with a large percentage of Fe^{3+} ions with higher intrinsic catalytic activity than Cr^{3+} . In this way, even though the amount Fe^{3+} in the catalyst is only 21 wt %, the activity of Fe^{3+} ions in this MIL-101 structure is similar or higher than when homometallic MIL-100(Fe) or MIL88-(Fe) are used as catalyst, due to a more impeded accessibility to the Fe^{3+} sites.

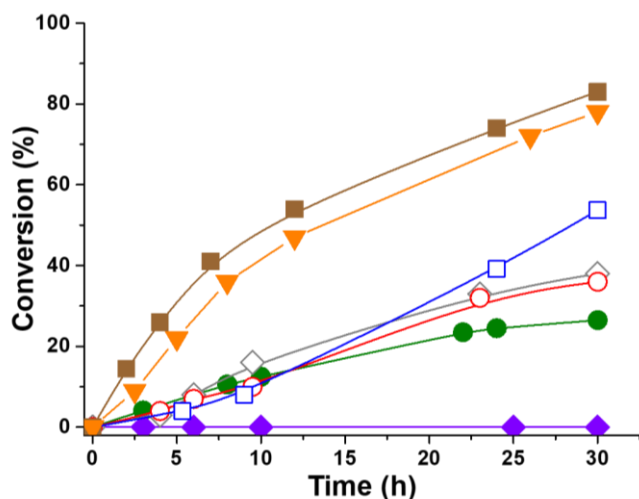


Figure 4. Catalytic activity for β -pinene using MIL-101(Cr/Fe) (4:1) (■), MIL-88(Fe) (◇), MIL-101(Fe) (○), MIL-101(Cr) (●), MIL-100(Fe) (▼), Fe^{3+} acetate (□) and in the absence of catalyst (◆). Reaction conditions: Catalyst (0.02 mmol of metal), β -pinene (800 μL , 5 mmol), formaldehyde (750 μL , 5 mmol), acetonitrile (1.26 mL), 80 $^{\circ}\text{C}$.

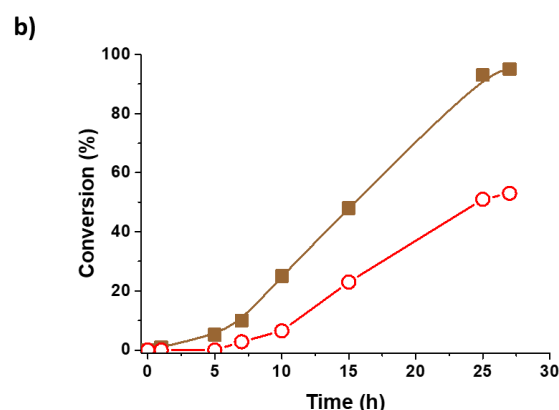
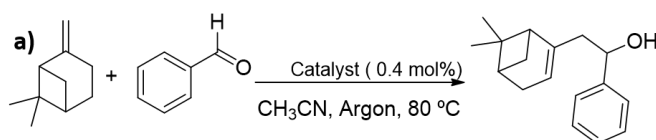


Figure 5. a) Prins addition of β -pinene and benzaldehyde. b) Temporal conversion plot of β -pinene using MIL-101(Cr/Fe) (4:1) (■) or MIL-100(Fe) (○) as catalyst. Reaction conditions: Catalyst (0.02 mmol of Cr+Fe), β -pinene (800 μL , 5 mmol), benzaldehyde (5 mmol), acetonitrile (1.26 mL), 80 $^{\circ}\text{C}$.

In agreement with the above proposal to rationalize the activity of mixed Cr/Fe MOF, the benefits of MIL-101(Cr/Fe) (4:1) respect to the MIL-100(Fe) as solid catalyst to promote the Prins reaction of β -pinene are even more clearly evidenced when a bulkier aldehyde such as benzaldehyde is used as substrate. Figure 5 shows a comparison of the time-conversion plot under the same conditions for the Prins reaction with benzaldehyde as reagent using MIL-101(Cr/Fe) (4:1) or MIL-100(Fe) as catalyst. As it can be seen the activity of MIL-101(Cr/Fe) (4:1) is higher than that of MIL-100(Fe) in spite of the 5-fold higher Fe^{3+} content of the latter.

To determine the influence of the Fe^{3+} content present in the mixed-metal MIL-101(Cr/Fe) solid on the catalytic activity for the Prins reaction, three MIL-101(Cr/Fe) containing different Fe^{3+} contents with values 2.5, 14 and 21 wt% (ratio Cr/Fe 36:1, 5.9:1, 4:1, respectively) were tested. Figure 6 presents the time-conversion plots for the Prins reaction promoted by these three mixed-metals MIL-101(Cr/Fe) solids. By plotting the initial reaction rate vs. the Fe^{3+} content it was observed that, although Cr^{3+} also catalyzes the reaction as it can be deduced from the initial reaction rate value at 0 Fe^{3+} content, the initial reaction rate increases with the Fe^{3+} content. This increase in the initial reaction rate means that Fe^{3+} is more active than Cr^{3+} to promote the reaction. Furthermore, an increase of TOF value at higher Fe^{3+} loadings was observed. In principle, the TOF value should be constant if all the Fe^{3+} sites were equally active. Therefore, the TOF increase along the Fe^{3+} percentage suggests some degree of cooperation among the Fe^{3+} sites.

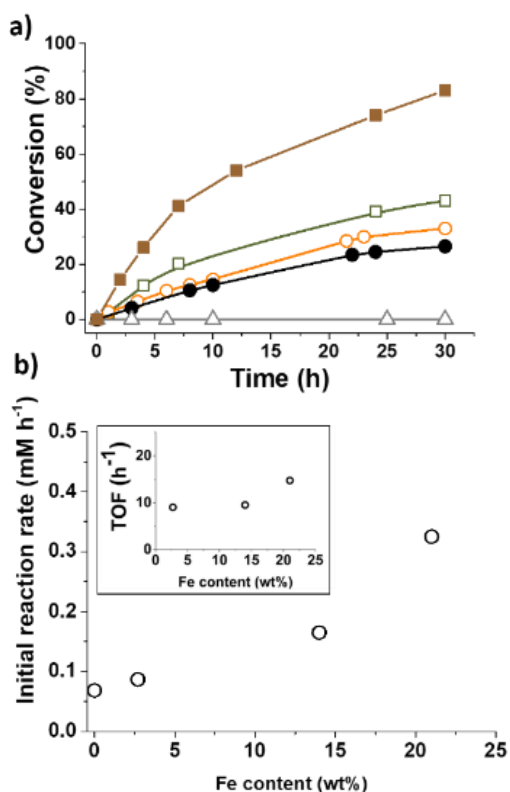


Figure 6. (a) Time-conversion plot for the Prins reaction between β -pinene and formaldehyde in the presence of MIL-101(Cr/Fe) (4:1) (\blacksquare), MIL-101(Cr/Fe) (5.9:1) (\square), MIL-101(Cr/Fe) (36:1) (\circ), MIL-101(Cr) (\bullet) and in the absence of catalyst (Δ). (b) Influence of the iron MIL-101 iron content on the initial reaction. The inset shows the influence of the iron content on the TOF; note that the activity of the MIL-101(Cr) has been subtracted to calculate the TON of mixed-metal MIL-101(Cr/Fe) solids. Reaction conditions: Catalyst (0.02 mmol of Fe+Cr), β -pinene (800 μ L, 5 mmol), formaldehyde (750 μ L, 5 mmol), acetonitrile (1.26 mL), 80 $^{\circ}$ C.

The catalytic activity of mixed-metal MIL-101(Cr/Fe) (4:1) was finally compared with that of homogeneous Cr^{3+} acetate and heterogeneous Cr_2O_3 or Fe_2O_3 solids (Figure 7). MIL-101(Cr/Fe) (4:1) exhibits the highest initial reaction rate in the series followed by the soluble Cr^{3+} salt. This observation is in agreement with previous results showing the benefits of metal ion isolation by immobilization in highly accessible metal clusters of porous MOFs acting as single-site catalysts.³⁷ On one hand, they exhibit similar or higher intrinsic activity to that of the homogeneous sites catalysts, and on the other, they undergo much slower deactivation by aggregation due to site immobilization.

As mentioned above and shown in Scheme 1, besides Nopol the formation of 10-pinen-3-ol was also observed in the presence of MIL-101. Selectivity values are presented in Figure 9. As described in previous reports, 10-pinen-3-ol arises from the allylic oxidation of β -pinene by molecular oxygen.^{38, 39} There are in the literature several studies showing that, besides as solid Lewis acid, MIL-101 can act also as oxidation

catalyst.^{14, 40} Under the present reaction conditions, the highest conversion and highest Nopol selectivity at final reaction time (30 h) was achieved using the homogeneous Cr^{3+} salt. Importantly, the use of the heterogeneous MIL-101(Cr/Fe) (4:1) catalyst results in a higher catalytic activity and higher Nopol selectivity respect to the metal oxides $\gamma\text{-Fe}_2\text{O}_3$ or $\alpha\text{-Cr}_2\text{O}_3$.

In the case of the mixed-metal MIL101(Cr/Fe) (4:1), the heterogeneity of the reaction was supported by performing a hot filtration test, observing that the reaction almost completely stopped after filtration of the solid (Figure 8). The minor conversion from 26 to 30 % from 4 to 30 h in the absence of solid after filtration of the catalyst was due to the occurrence of β -pinene oxidation to 10-pinen-3-ol, but not to the β -pinene addition to formaldehyde.

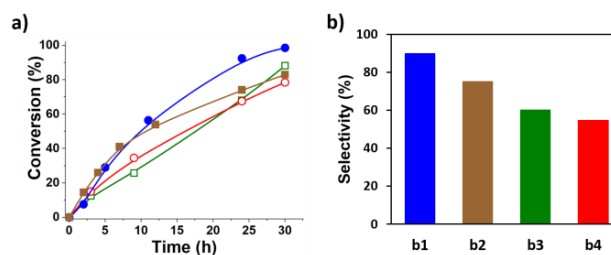


Figure 7. (a): Time-conversion plots of β -pinene using MIL-101(Fe/Cr) (4:1) (\blacksquare), Cr^{3+} acetate hydroxyde (\bullet), Fe_2O_3 (\square) or Cr_2O_3 (\circ) as catalysts and (b): Nopol selectivity for Cr^{3+} acetate hydroxyde (**b1**), MIL-101(Fe/Cr) (4:1) (**b2**), Fe_2O_3 (**b3**) or Cr_2O_3 (**b4**) as catalysts. Reaction conditions: Catalyst (0.02 mmol metal), β -pinene (800 μ L, 5 mmol), formaldehyde (750 μ L, 5 mmol), CH_3CN (1.15 mL), 80 $^{\circ}$ C.

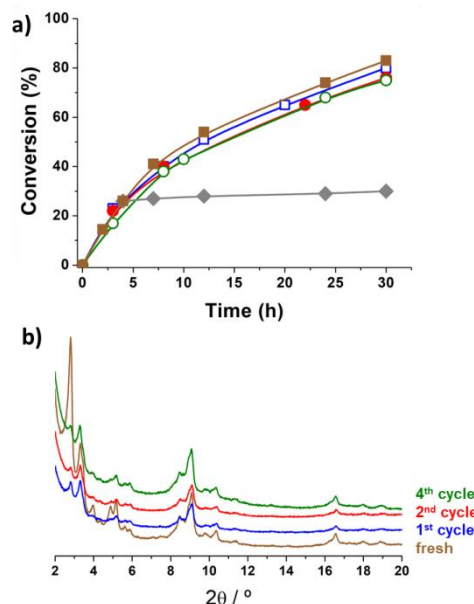


Figure 8. (a) Hot filtration and reusability tests 1st use (\blacksquare), 2nd use (\square), 3rd use (\bullet), 4th use (\circ) and without catalyst (\blacklozenge). (b) PXRD patterns for the fresh and used MIL-101(Cr/Fe) (4:1). Reaction conditions: Catalyst (0.02 mmol metal), β -pinene (800 μ L, 5 mmol), formaldehyde (750 μ L, 5 mmol), CH_3CN (1.15 mL), 80 $^{\circ}$ C.

Catalyst stability was finally studied by performing four consecutive runs using the same MIL-101(Cr/Fe) (4:1) sample. Comparison of the temporal profiles of first and consecutive reactions showed only a slight decrease of the catalytic activity, while maintaining the same product selectivity of about 75 %. This partial deactivation is not due to degradation of the MIL-101(Cr/Fe) (4:1) catalyst, since little changes were observed in PXRD patterns recorded after each run compared to the fresh material (Figure 8(b)). A loss of intensity of the first Bragg reflexions was observed, likely due to remaining reactants in the pores. Analysis by ICP-OES of the chromium and iron content in the liquid reaction phase after filtration of the solid catalyst showed that the concentration of these metals in the liquid phase corresponds to about 0.002 and 0.02 wt% of the initial chromium and iron content of the fresh catalyst, respectively, in agreement with the absence of degradation of the MIL-101(Cr/Fe). Thus, the slight deactivation observed might be attributed to adsorbed reaction products or by-products partially blocking the pores and/or active sites, therefore, causing an apparent decrease on the performance of the material.

Characterization of the acid sites

In order to get some insight on the higher catalytic activity of the mixed-metal MIL-101(Cr/Fe) respect to MIL-101(Cr), an *in situ* infrared spectroscopy study was undertaken to investigate the Lewis acid sites on those solids.⁴¹ Following previous works using MIL-100(Fe)^{42, 43} or MIL-100(Cr)^{18, 44} materials, the MIL-101 samples under study were activated at 150 °C under vacuum for 3 h to remove the water molecules coordinated to the metal centers. Then, the acid strength of the resulting coordinatively unsaturated sites (CUS) of the MIL-101 solids was studied using CD₃CN and CO as probe molecules.

In the case of CD₃CN adsorption on activated MIL-101(Cr/Fe) (4:1), the characteristic $\nu(\text{CN})$ and $\nu(\text{CD}_3)$ bands due to the presence of CD₃CN on both the CUS or physisorbed in the solid were recorded (Figure 9). The broad $\nu(\text{CN})$ band with the maximum appearing at 2320 cm⁻¹ can be assigned to the coordinated deuterated acetonitrile on the Cr³⁺ and/or Fe³⁺ Lewis sites present in the MIL-101(Cr/Fe) (4:1) solid. This $\nu(\text{CN})$ wavenumber is slightly lower than that found for MIL-100(Al) (2321 cm⁻¹), a fact that agrees with the polarizing power of Al³⁺>Fe³⁺>Cr³⁺.⁴⁵ In the case of MIL-101(Cr/Fe) (4:1) solid an additional $\nu(\text{CN})$ band at higher wavenumber (2350 cm⁻¹) appeared and could be assigned to adsorption of CD₃CN on metal defects with a higher Lewis acidity as it has been proposed in the case of the MIL-100(Al) ($\nu(\text{CN})$ 2341 cm⁻¹).⁴⁵ Physisorbed acetonitrile is characterized by the $\nu(\text{CN})$ band appearing at lower wavenumbers (2274–2261 cm⁻¹). The band corresponding to the $\nu(\text{CD}_3)$ can be observed at 2112 cm⁻¹. For comparison, CD₃CN adsorption on activated MIL-101(Cr) was also studied. The higher $\nu(\text{CN})$ wavenumbers of MIL-101(Cr/Fe) (4:1) (2320 and 2350 cm⁻¹) vs. MIL-101(Cr) (2318 and 2344 cm⁻¹) is an indication of the stronger Lewis acidity of the former. Again, this observation agrees with the charge to radius ratio

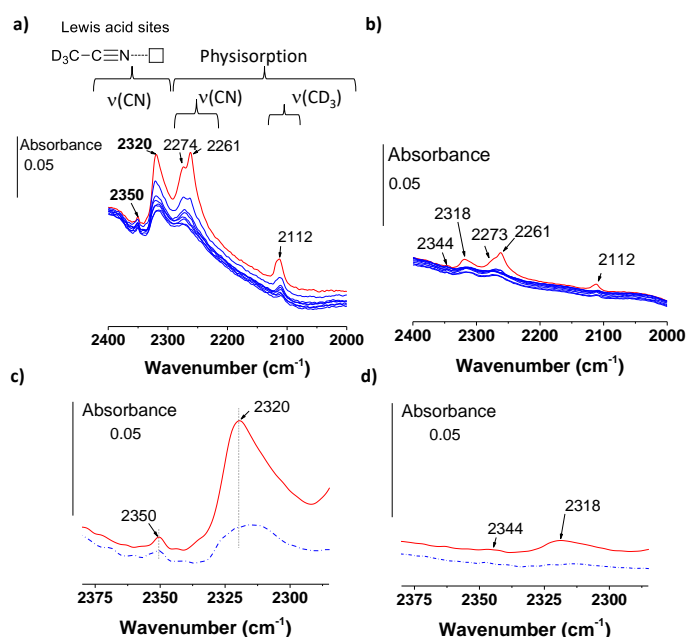


Figure 9. FT-IR spectra of CD₃CN adsorbed in MIL-101(Cr/Fe) (a, c) or MIL-101(Cr) (b, d). Red line spectra show the equilibrated CD₃CN adsorbed in MOFs; Blue line spectra show the retained CD₃CN after outgassing the corresponding MOF for 1 h period.

of the Cr³⁺ vs. Fe³⁺ and, then, with the expected higher polarizing power of Fe³⁺ vs. Cr³⁺ as Lewis center.⁴⁵ Once the CD₃CN equilibrium was reached for MIL-101(Cr/Fe), outgassing the samples led to a decrease of the area under the peak at 2350 cm⁻¹ in lesser extent than that of 2320 cm⁻¹. This suggests that the higher the Lewis acid strength of the metal nodes, the higher the observed $\nu(\text{CN})$ wavenumber and, therefore, the stronger the adsorption of the probe.

Importantly, for the same weight of MIL-101(Cr/Fe) and MIL-101(Cr) solids the CD₃CN adsorption measurements indicated a ten-fold increase of the area corresponding to the peaks at 2350 and 2320 cm⁻¹ in MIL-101(Cr/Fe) (4:1) compared to that of MIL-101(Cr). The higher $\nu(\text{CN})$ wavenumber value and the higher peak area in the case of MIL-101(Cr/Fe) respect to MIL-101(Cr), indicates the presence of higher density and stronger Lewis acidity of the metal nodes in MIL-101(Cr/Fe) respect to MIL-101(Cr). These results are in sharp contrast with those previously reported for MIL-100(Cr) vs. MIL-100(Fe) that indicate a higher Lewis acid density for MIL-100(Cr).⁴⁵ This might be a consequence of the partial poisoning of the strongest Lewis sites in MIL-100(Fe) and/or the presence of structural defects in MIL-100(Cr). In any case, it seems that for the mixed metal Cr/Fe solid, the presence of the two metals is associated with a higher density of structural defects characterized by the 2350 cm⁻¹ $\nu(\text{CN})$ band and stronger CUS Lewis sites corresponding to the 2320 cm⁻¹ $\nu(\text{CN})$ band.

The acidity of the MIL-101(Cr/Fe) (4:1) material was further characterized by IR spectroscopy by using CO as probe molecule. Figure SI-10 shows three main $\nu(\text{CO})$ bands corresponding to CO coordinated with Lewis acid sites [$\nu(\text{CO})$ at 2195 cm⁻¹], CO interacting with Brønsted acid sites [$\nu(\text{CO})$ at 2161 cm⁻¹] and physisorbed species [$\nu(\text{CO})$ at 2136 cm⁻¹]. In the

case of MIL-101(Cr), similar CO bands were also observed although slightly shifted to lower wavenumber values for the Lewis and Brønsted acid sites respect to the mixed-metal MIL-101(Cr/Fe) (4:1). These observations are in agreement with the higher acidity of the MIL-101(Cr/Fe) (4:1) solid in respect to MIL-101(Cr) observed using CD₃CN as probe. Once the CO equilibrium was achieved in each MIL-101 sample, a vacuum treatment slightly decreased the peak intensity of the Lewis centers, while the signals of the weaker Brønsted acid sites as well as those of physisorbed CO species were almost completely vanished.

This IR spectroscopy study allowed us concluding that the higher catalytic activity observed for MIL-101(Cr/Fe) arises mainly from the increase of Lewis acidity due to the incorporation of Fe ions at the metal nodes. In agreement with this conclusion, when the Prins reaction was carried out using the MIL-101(Cr/Fe) (4:1) solid adding pyridine as a base,⁴⁶ the resulting catalytic activity was drastically reduced (Figure SI-9 in supporting information).

Conclusions

The present study illustrates the advantages of combining the structural robustness and porosity, characteristic of MIL-101(Cr) material with the presence of Fe³⁺ as active Lewis sites. Overall, the present study exemplifies the potential that mixed-metal MOF offer in catalysis, particularly when preparation methods make possible to prepare crystalline materials with a wide range of metal proportions as it is the case of the present MIL-101 materials.

Experimental section

Materials

All the reagents and solvents employed in this work were of analytical or HPLC grade and supplied by Sigma-Aldrich.

Catalyst preparation

MIL-101(Cr),^{10, 12} MIL-101(Fe)¹⁴ and MIL-88B(Fe)⁴⁷ have been prepared as previously reported. MIL-101(Cr/Fe) materials were prepared by mixing 1.5 mmol terephthalic acid, 1 mmol metal source (Cr and Fe) and 10 mL tetramethylammonium hydroxide (TMAOH) solution (0.06 M) in a Teflon reactor. More precisely, 0.75 mmol Cr(NO₃)₃·9H₂O and 0.25 mmol of Fe⁰ were used for MIL-101(Cr/Fe) (4:1) and 0.85 mmol Cr(NO₃)₃·9H₂O and 0.15 mmol of Fe⁰ for MIL-101(Cr/Fe) (5.9:1). The reaction mixtures were heated at 150 °C for 48 h under autogenous pressure. The obtained solids were isolated by centrifugation (20 min, 14500 rpm) and washed three times with water and three times with abs. EtOH. Density separation was performed to remove any iron oxide impurities from the MOFs. In order to remove the residual ligand molecules from the pores, the solids were treated with a solution of KF (0.1 M) for 1 h. The purified products were dried at 100 °C overnight, prior to their characterization.

Characterization

Powder X-ray diffraction patterns were obtained on a Siemens D5000 diffractometer using Cu K_α 1,2 radiation (λ=1.5406Å).

N₂ sorption measurements were performed at 77 K on a BELL Japan Belsorp Max apparatus, after an overnight activation of the samples at 150 °C, under primary vacuum.

⁵⁷Fe Mössbauer spectra were recorded at 300 and 77 K using a ⁵⁷Co/Rh γ-ray source moving with a triangular velocity form and mounted on a conventional electromagnetic drive. The hyperfine structure was modeled by a least-square fitting procedure involving quadrupolar doublets composed of Lorentzian lines. The isomer shift values (IS) are referred to α-Fe at 300 K, while the samples consist of a thin layer of powder containing about 15 mg of Fe.

The metal content (Cr or Fe) of the MOF samples was determined by using inductively coupled plasma optical emission spectroscopy (ICP-OES, Perkin-Elmer). Previously to ICP-OES analysis, the solid samples (10 mg) were digested using concentrated HNO₃ (65 wt%; 20 mL) at 80 °C for 24 h. The resulting solution was diluted with Milli-Q water and, then, analyzed by ICP-OES.

FTIR spectroscopy was performed by compression the solid samples at 2 Ton in air prior to recording the spectra. In some cases the sample was mixed with SiO₂ as binder. Then, the samples were placed in a quartz cell equipped with KR5-5 (thallium bromoiodide) windows. The cell was connected to a vacuum line for activation ($P_{residual} = 10^{-5}$ mbar for 2 h) prior to CO or CD₃CN adsorption. For the CO adsorption experiments, the temperature of the pellet was decreased to about -170 °C by cooling the sample holder with liquid N₂. CD₃CN adsorption experiments have been carried out at room temperature. The addition of CO in the cell was from 0 to 2.56 mbar, approximately. CD₃CN adsorption was performed at 0.26 or 0.82/0.35 mbar for MIL-101(Cr/Fe) or MIL-101(Cr), respectively. Transmission IR spectra were recorded in the 4000-400 cm⁻¹ range, at 4 cm⁻¹ resolution, on a VERTEX 70 spectrophotometer equipped with DTGS (Deuterated triglycine sulfate detector).

Catalytic reactions

Prins reaction. Prior to its use as catalyst, the corresponding amount of the solid (typically 0.02 mmol of Cr+Fe metal) was dried in a two-necked round bottom flask (25 mL) at 100 °C for 12 h. Then, the system was heated at 150 °C under vacuum for 16 h. Subsequently, the system was cool down to 80 °C and the reagents added to the system under continuous magnetic stirring in the following order: acetonitrile (1.26 mL), formaldehyde (750 μL, 5 mmol) and β-pinene (800 μL, 5 mmol).

Selective quenching experiments were carried out as described above, but adding pyridine at 20 mol % respect to the substrate.

Reuse experiments were carried out under the general reaction conditions. At the end of each run, the catalyst was recovered by filtration through a Nylon membrane (0.2 μm). The retained solid was washed by flushing hot acetonitrile (50 °C, 400 mL) and, then, ethanol (20 °C, 600 mL). The resulting

solid was dried in an oven at 100 °C for at least 12 h before being employed in a subsequent catalytic cycle.

The progress of the reaction was followed by gas chromatography (GC) using a flame ionization detector (FID). During the reaction several aliquots (100 µL) were sampled, then diluted with acetonitrile (0.5 mL) and filtered using syringe filters (0.2 µm Nylon). Nitrobenzene was finally added as external standard (10 µL) to each reaction aliquot. Quantification was carried out using a calibration plot for each commercially available compound. The samples were analyzed immediately after being taken from the reaction mixture.

Conflicts of interest

There are no conflicts to declare

Author contribution

Materials were synthesized by Effrosyni Gkaniatsou, Mónica Giménez-Marqués, Nathalie Steunou and Clémence Sicard. Catalytic experiments were initially performed by Andrea Santiago-Portillo and finalized by Cristina Vallés-García under the supervision of Sergio Navalón and Mercedes Álvaro. Mossbauer spectra were recorded and interpreted by Jean-Marc Grenèche. The work was conceived by Clémence Sicard, Sergio Navalón, Christian Serre and Hermenegildo García.

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