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Tuning Cellular Biological Functions Through the Controlled Release of NO from a Porous Ti-MOF

Rosana V. Pinto⁺, Sujing Wang⁺, Sergio R. Tavares, João Pires, Fernando Antunes, Alexandre Vimont, Guillaume Clet, Marco Daturi, Guillaume Maurin, Christian Serre,* and Moisés L. Pinto*

Abstract: Materials for the controlled release of nitric oxide (NO) are of interest for therapeutic applications. However, to date, many suffer from toxicity and stability issues, as well as poor performance. Herein, we propose a new NO adsorption/release mechanism through the formation of nitrites on the skeleton of a titanium-based metal-organic framework (MOF) that we named MIP-177, featuring a suitable set of properties for such an application: (i) high NO storage capacity ($3 \mu\text{mol mg}^{-1} \text{solid}$), (ii) excellent biocompatibility at therapeutic relevant concentrations (no cytotoxicity at $90 \mu\text{g mL}^{-1}$ for wound healing) due to its high stability in biological media ($< 9\%$ degradation in 72 hours) and (iii) slow NO release in biological media (≈ 2 hours for 90% release). The prospective application of MIP-177 is demonstrated through NO-driven control of mitochondrial respiration in cells and stimulation of cell migration, paving the way for the design of new NO delivery systems for wound healing therapy.

Introduction

Nitric oxide (NO) plays an important role in signaling pathways in human physiological and pathological processes, garnering considerable interest in its potential as a therapeutic agent. In particular, exogenous delivery of NO has been widely demonstrated to have a therapeutic effect in wound healing^[1,2] through its promotion of vascular permeability^[3] and killing of pathogens.^[4] It is also currently considered as an anti-oncogenic agent.^[5] However, due to the high toxicity and short half-life, delivery of NO to the target sites in a controlled manner remains a great challenge. For this purpose, ordered nanoporous materials have recently been proposed as carriers

for the controlled local delivery of pure NO that homogeneous molecular donors fail to achieve.^[6] These nanoporous delivery materials, including zeolites,^[7,8] titanosilicates,^[9,10] and metal-organic frameworks (MOFs),^[11–14] have demonstrated considerable NO storage and release capacity. Some of these materials can even form metal-NO coordination bonds pointing towards the pores, allowing the adsorption of large amounts of NO with a better controlled release. Another approach, using porous coordination polymers (PCPs) functionalized with N-nitrosamine functional groups or incorporating nitro-containing imidazolate ligands with poor reactivity into the PCP frameworks, demonstrated efficient and controlled release of NO using light radiation.^[15,16]

However, the toxicity and stability issues of all these nanomaterials, as well as the biological impact of the NO encapsulation/release have scarcely been explored so far.^[11,12,17,18] Moreover, the design of the ideal nanoporous NO carrier is still a very demanding challenge since the material needs to be biocompatible and stable in aqueous biological media (i.e., no signs of toxicity at therapeutic concentrations or release of metals when in contact with cells for several days), feature strong NO storage capacity (currently best materials are capable of storing 1–7 $\mu\text{mol NO mg}^{-1}$),^[12,13,19] and slow NO release in biological media during several hours since the biological activity and toxicity of this molecule is concentration-dependent.^[20–22] To our knowledge, such a critical combination of features has not been achieved so far. The development of new solid NO donors is thus required to overcome the limitations of the state-of-the-art molecular donors, that are, inadequate long-term delivery,

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instability during storage and issues associated to systemic circulation of potentially toxic nanovectors.^[6,23] To that end, biocompatible MOFs with coordinatively unsaturated sites (CUSs) in their porous structures could provide a potential solution.

MOFs are one of the newest classes of porous crystalline materials that have recently been explored for a wide range of medical applications,^[24–27] with a great potential for the encapsulation and release of drugs.^[25] Nevertheless, biocompatible MOFs considered so far have exhibited major drawbacks, such as limited structural stability in biological media strongly affecting the NO release performance, also leading to the release of metals into the media which may, in some cases, induce toxicity.^[13,26,28] Indeed, the MOF structure has to be sufficiently stable, particularly if an effective release over several hours is targeted as, for example, in wound healing applications.

Herein, we have considered a very recently reported titanium carboxylate MOF, denoted MIP-177 (Figure 1), in its low temperature, formate-free form (MIP stands for the materials of the Institute of Porous Materials of Paris), that fulfils the requirements for an ideal NO delivery system. ■ ■reference required for MIP-177? ■ ■ We here demonstrate that this non-toxic material is highly stable in biological media and exhibits a porous structure featuring titanium metal sites on its inorganic building unit in an ideal position to coordinate NO by the uncommon in situ formation of complexed nitrites. The performance of this new donor in biological systems has been analyzed by testing different NO concentration ranges until the desired effect is observed, that is, the stimulation of cell migration. This is a fundamental process in wound healing and, to the best of our knowledge, is being explored for the first time with this class of NO donors. This MOF material, which offers the ability to exercise control over certain basic biological processes, represents a step forward in the development of new strategies for therapeutic uses such as the acceleration of wound healing.

Results and Discussion

Synthesis and Characterization of the Ti-MOF

The use of active Ti open metal sites was reported to be attractive for NO coordination and delivery in the microporous titanosilicate ETS-4.^[9] However, harsh preparation conditions required on a small scale and limited accessible porosity resulting in restricted accessibility to the Ti metal sites are clear limitations of this solid. Moreover, ETS-4 exhibits fast release of NO upon contact with biological media due to its high hydrophilicity, as will be shown here. To address those limitations, MIP-177 in its formate-free form with a chemical formula $Ti_{12}O_{15}(mdip)_3(OH)_6(H_2O)_6$ ($mdip = 3,3',5,5'$ -tetracarboxydiphenylmethane), has been selected as NO storage and release candidate based on its following features: (i) robust porosity with adequate adsorption sites to enable large NO uptake and release. As shown in Figure 1, MIP-177 comprises nanosized one-dimensional (1D) channels with a free diameter of 1.1 nm running along the *c*-axis. Each $Ti_{12}O_{15}$ secondary building unit (SBU) possesses three pairs of terminal OH/ H_2O species facing the pore, which can be removed thermally to generate bridged Ti–OH–Ti as evidenced by coupling density functional theory (DFT) calculations and in situ infrared spectroscopy (FTIR) analysis (see Supporting Information); (ii) excellent hydrolytic and chemical stability under various conditions as reported previously^[29] and more specifically in biological media. This is of great importance to avoid any toxicity and performance fluctuation during application in real biological systems; and (iii) an easy and environmentally friendly preparation for future scale-up applications. MIP-177 particles are of elongated shape with uniform sizes between 0.1 and 0.2 μm in length that form agglomerates of about 1 μm (Figure S21, Supporting Information), which are biocompatible as shown below.

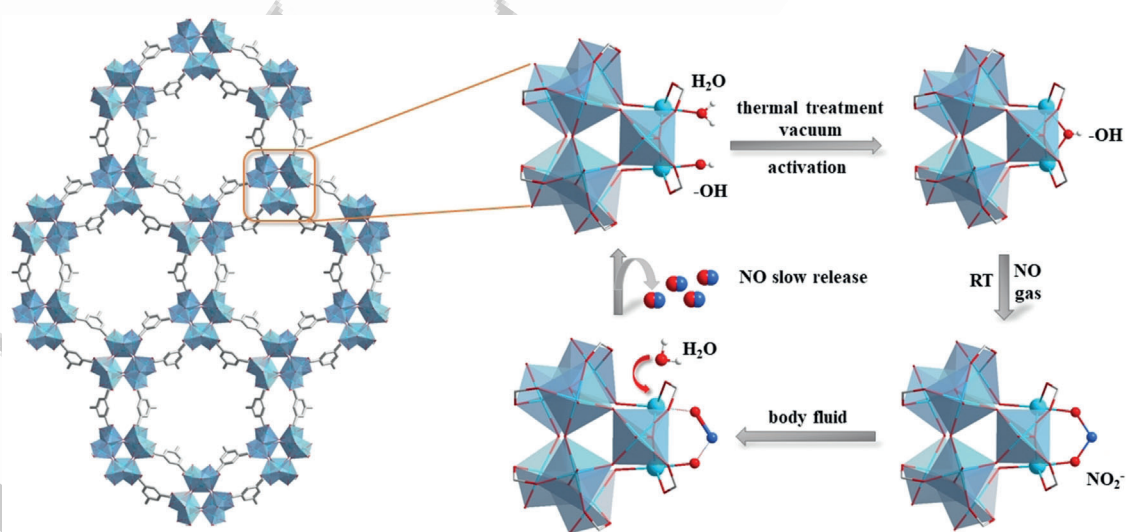


Figure 1. MIP-177 microporosity and binding/release mechanism: viewed along the *c*-axis (left), adsorption and controlled release cycle of NO under the tested conditions (right) is shown.



Computational Prediction of Interactions Between NO and MIP-177

DFT calculations were first performed for the “empty” MOF material; the formate groups connecting two Ti sites were replaced by one terminal OH bound to one Ti site, leaving the second Ti site uncoordinated to mimic a dehydrated state. The DFT-geometry optimization of such a configuration led to the formation of bridging OH groups bonded to both Ti sites (Figure S1), which is in agreement with FTIR observations. These OH groups are the preferential interacting site for NO with a H(OH)–O(NO) distance of 2.9 Å (Figure S2a), associated with a high binding energy of 72 kJ mol⁻¹ which however remains lower than the value previously reported for the adsorption of NO on the Ni CUSs in Ni-CPO-27 (90 kJ mol⁻¹).^[30] We have evidenced that even when OH is forcibly bound in the terminal position to one Ti site, leaving the neighbouring Ti site uncoordinated, NO does not interact strongly with the Ti CUS (Figure S2a). Furthermore, the comparison of the total energies of both configurations (Figure S2a and S2b) indicates that the interaction of NO with the OH bridge is highly favoured. Therefore, as we will demonstrate in the subsequent sections, the bridging OH group will react with NO to produce bridging or chelating nitrites (Figure 1).

Therapeutic NO Release

NO adsorption on MIP-177 reveals a maximum storage capacity of 9% (w/w) (total of 3 μmol mg⁻¹_{solid}) after 27 hours (Figure 3a), with fast loading kinetics, reaching 6% (w/w) in the first 5 hours. Considering the theoretical number of binding sites available in MIP-177 and the in situ formation of

nitrite species to generate Ti₁₂O₁₅(mdip)₃(NO(OH))₆ (Figure 1 and in situ IR discussion below), the maximum expected loading capacity is about 9.6 wt% (see Supporting Information for details), which is close to the observed maximum storage capacity. This indicates that MIP-177 is among the best MOFs so far reported in the literature (0.8–7 μmol mg⁻¹) for NO storage.^[12,13,31,32]

The kinetic desorption curve in Figure 3b) shows quick NO release in the first 2 hours, reaching an equilibrium after 20 hours. NO amounts released from MIP-177 (1.7 μmol mg⁻¹; about 55% of the loaded amount) exceed those of previously reported porous materials. For example, zeolites, clay-based materials and titanosilicates such as ETS-10, ETS-4 and some modified formulae all present values ≤ 1.7 μmol mg⁻¹.^[9,10,33–36] Furthermore, the release performance of MIP-177 is higher than reported for most MOFs, only CPO-27 exceeds the value of MIP-177.^[12,13,19] Remarkably, 45% of the loaded NO is retained in the material even under high vacuum, which is in agreement with the formation of strongly-bound nitrite species (Figure 1 and 2). The retained NO is released when the material is introduced to biological media (demonstrated below).

NO adsorption and desorption were also monitored by FTIR (Figure 2). Several species were formed over a 24-hour adsorption period, which subsequently exhibited different desorption behaviors. Shortly after NO introduction, a new band quickly developed at 1850 cm⁻¹ and was assigned to weakly adsorbed NO, as previously reported for NO adsorption on MIL-88(Fe).^[37] This band did not evolve significantly after 3 hours and the weak interaction with the pore walls was confirmed by its disappearance upon evacuation at room temperature. At the same time, a band at 1210 cm⁻¹ appeared reaching its maximum after about 3 hours, at which point it started decreasing (Figure S5). This spectral range is typical of

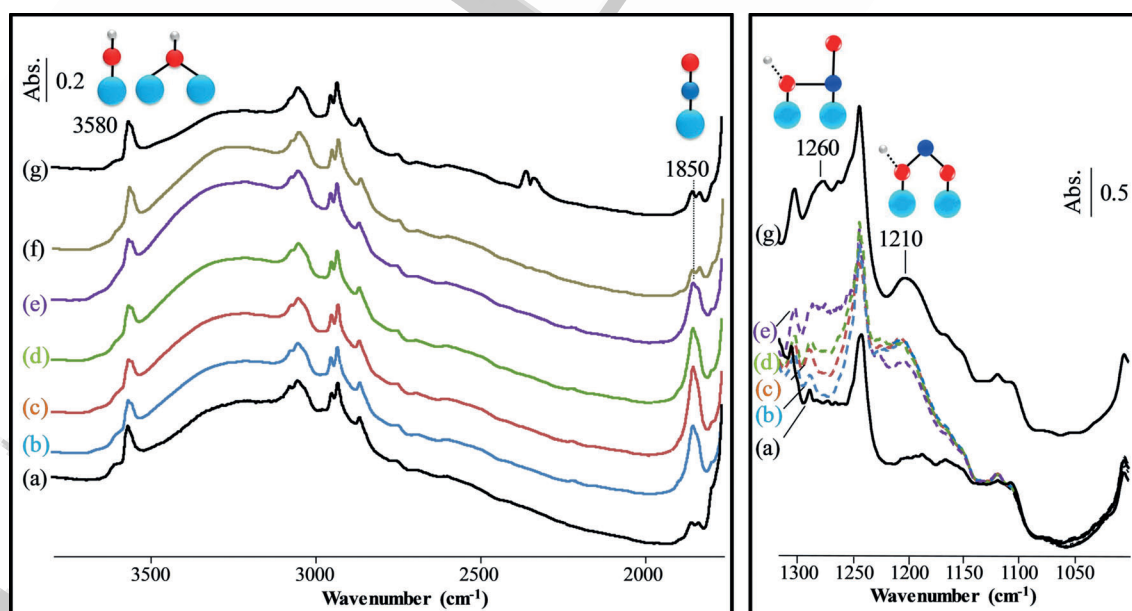


Figure 2. Infrared absorption spectra of MIP-177 recorded after thermal treatment at 120°C (a) and after introduction of NO at room temperature ($p_{\text{NO}}=80$ kPa) at exposure times of 20 min (b), 75 min (c), 195 min (d), 20 hours (e), followed by outgassing at room temperature for 20 min (f) and at 120°C for 2 hours (g). Adsorbed species representation: ● = Ti, ● = N, ● = O, ● = H.



bridging nitrites or nitrates.^[38] The possibility of nitrate formation can be eliminated as the parallel development of characteristic bands around 1020 cm^{-1} is not observed in the spectrum. Therefore, the band at 1210 cm^{-1} can be confidently assigned to bridging nitrites $(\text{Ti-O})_2\text{=N}$. These modifications are associated with an initial decrease of the band at 3580 cm^{-1} and the development of a broad band around 3300 cm^{-1} implying the involvement of the hydroxyl species and its perturbation (or even shift, in the form of adsorbed water molecules) in this first step, to allow nitrite formation. Adsorption proceeds through a second step since another band develops at 1260 cm^{-1} , which was assigned to chelating nitrites.^[38,39] Unlike the previous bands, this one slowly increased over the first 200 minutes, and continued growing steadily for up to 24 hours. A slight development of the band at ca. 3300 cm^{-1} was also observed, while the bands at 1210 and 3580 cm^{-1} were progressively decreasing in parallel, with a trend similar to that of the band at 1260 cm^{-1} (Figure S5). This is a clear indication that chelating nitrites form at the expense of some bridging nitrites and provoke the removal of weakly coordinated hydroxyls, as schematized in Figure S6. Upon evacuation, bands at both 1210 and 1260 cm^{-1} persisted, conversely to the band at 1850 cm^{-1} , evidencing their stability. Therefore, we conclude that NO adsorption induces the formation of both weakly adsorbed species and stable nitrites.

This is in excellent agreement with the release of only ca. 55 % of the loaded NO after outgassing, leaving the remaining 45 % of the loaded NO in the MOF as very stable nitrites bonded to a pair of Ti sites.

To further investigate chelating and bridging nitrite/nitrate adsorptions, DFT calculations were performed (see Figure S3 and S4). In case of the nitrite adsorption, the DFT calculations showed that chelating and bridging species may coexist since the two geometries converge towards a very similar energy. The configuration corresponding to the bridging nitrate species shows a substantially higher stability than that of the chelating nitrate species, with a relatively high energy difference of 17.1 kJ mol^{-1} . Finally, regarding the formation of the nitrite/nitrate species, the energy comparison between corresponding configurations (see discussion in the Supporting Information) indicated that the nitrite adsorption is much more favored. We estimated a binding energy of 134.1 kJ mol^{-1} for the nitrites on the Ti sites (Supporting Information) which is significantly higher than the values previously reported for Ni-CPO-27.^[30] Therefore, the DFT simulations and the IR measurements concur in terms of the presence of stable nitrite species chemisorbed on the Ti sites, which support the adsorption results (Figures 3a and b).

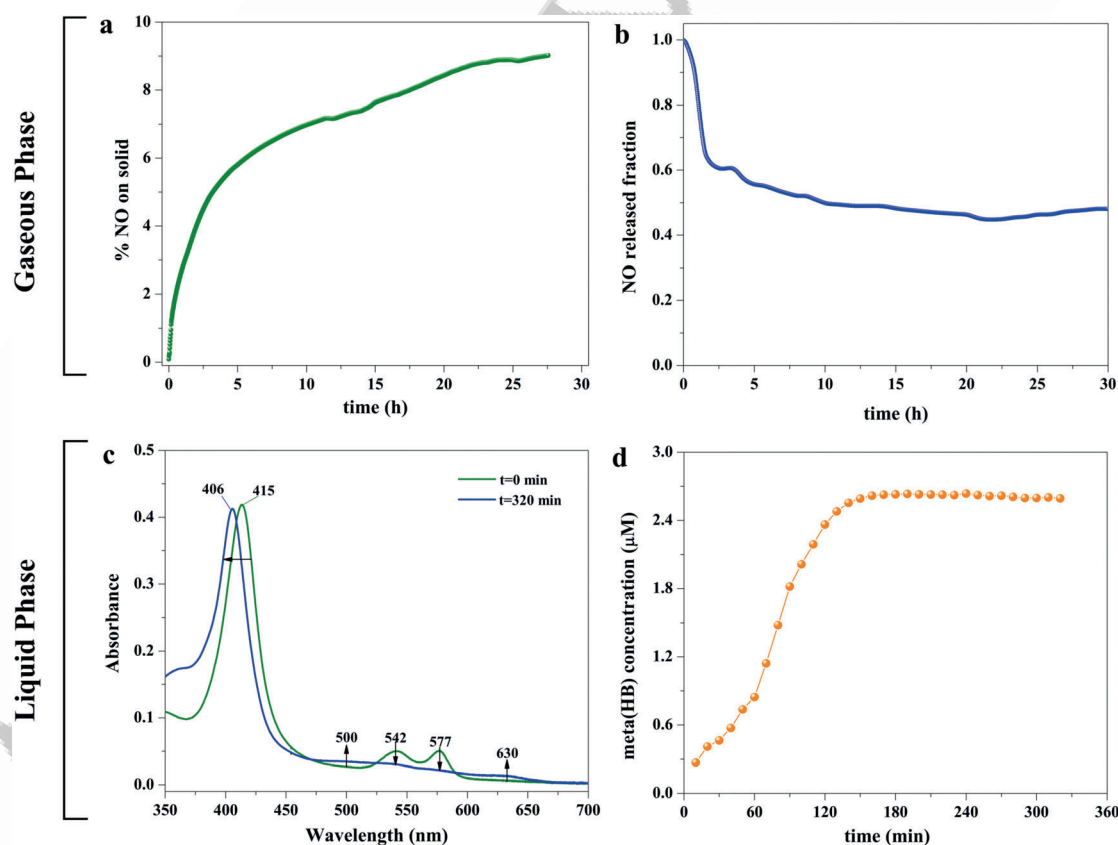


Figure 3. NO adsorption/release capacities in gas and liquid phase. a, b) NO kinetic profiles performed on a gravimetric apparatus, at $25\text{ }^{\circ}\text{C}$: a) represents NO adsorption performed at 80 kPa and b) NO release under vacuum of the studied material; c, d) NO release in liquid phase using the oxyhemoglobin assay, performed at $25\text{ }^{\circ}\text{C}$ using phosphate buffer solution: c) comparison between the initial UV/Vis spectrum of the oxyhemoglobin solution with the final UV/Vis spectrum (after 3 h) and d) NO kinetic release profile from the MIP-177 loaded with NO, in oxyhemoglobin solution.

1 These experimental and computational results shed light
2 on a new conceptual mechanism of NO storage/release using
3 porous solids. Until now, only MOFs bearing single unsatu-
4 rated metal sites (CUS), such as CPO-27/MOF-74 or MIL-
5 100(Fe), have been considered for NO delivery applications.
6 The adsorption/release mechanism relied on direct NO
7 coordination onto these metal sites. Herein, MIP-177, with
8 its unusual building unit involving two neighboring Ti sites,
9 allows the in situ formation of bridging or chelating nitrite
10 species upon adsorption of NO, resulting in a species more
11 strongly bound to the framework than a single-bonded NO
12 molecule to a typical CUS.

13 Previously, NO-loaded MOFs were mostly studied for
14 antimicrobial or anti-thrombogenic applications, in most
15 cases not in direct contact with biological fluids. These media
16 are more challenging for the controlled release of NO due the
17 faster exchange between NO and liquid water as well as the
18 presence of phosphates that rapidly degrade most metal
19 carboxylate MOFs. Given the enhanced stability of NO
20 species bound to MIP-177, we decided to consider this new
21 NO donor for biological applications. NO release kinetics
22 were thus further evaluated in the liquid phase, utilizing
23 several complementary methods. Using an oxyhemoglobin
24 assay^[40] (Figure 3c and d), we revealed the selective reaction
25 of NO with oxyhemoglobin when NO-loaded MIP-177 was
26 placed in contact with the solution. Changes in the UV/vis
27 spectrum over time (Figure 3c) were attributed to the trans-
28 formation of oxyhemoglobin into methemoglobin, demon-
29 strating NO release (this method is not sensitive to nitrite or
30 nitrate species). The NO release kinetics (Figure 3d) show
31 a stabilization in the concentration after about 150 minutes,
32 indicating complete release or complete oxyhemoglobin
33 consumption. Nevertheless, the initial slope shows a slow
34 release of NO. In another test, the loaded material was
35 introduced to RPMI-1640 cell culture medium at ambient
36 conditions. NO release was followed by a direct electro-
37 chemical measurement (NO selective electrode). For a MIP-
38 177 concentration of 450 $\mu\text{g mL}^{-1}$, the release reached a max-
39 imum of 100 nM (concentration decreases afterwards as NO
40 can react with O_2 in the medium) with a very attractive slow
41 kinetic profile, showing a controlled release over approx-
42 imately 2 hours, which has not been observed for any other
43 porous material before (Figure S7 and discussion/comparison
44 with other materials). Such a slow release is likely related to
45 the formation of bridging or chelating nitrites inside the MIP-
46 177, which are decomposing slowly through a reaction with
47 water and rehydration of the Ti metal sites, providing a stable
48 concentration of NO in the liquid phase^[41,42] (NB: ion
49 exchange with phosphates is less likely, since no significant
50 phosphate concentration was found in MIP-177 after some
51 days in phosphate solution). A Griess assay was also
52 performed for 2 hours using different concentrations of NO-
53 loaded MIP-177 under in vitro conditions (Figure S8). The
54 overall results obtained for NO release in the liquid phase
55 demonstrated that MIP-177 can release active NO with
56 particularly slow release kinetics (an extended discussion of
57 this point is presented in the Supporting Information). These
58 unique properties can be explained by the different nature of
59 NO bonding inside the pores of MIP-177 compared to that of

1 other MOFs. Indeed, this type of binding is more similar to
2 NO chemisorption in some polymers and dendrimers to form
3 RSNOs or NONOates.^[20] However, MIP-177 shows the
4 advantage of a larger storage capacity, affords more control
5 over the molecular structure of the NO vehicle, has very low
6 toxicity and provides additional optionality via its large pores
7 which may be used to store additional drug molecules for
8 combined synergic delivery and biological effect.^[20,26]

9 A toxicity assessment using two different cell lines (HeLa
10 and primary human epidermal keratinocytes—HEK293T) was
11 performed with MIP-177 at 180 $\mu\text{g mL}^{-1}$ and 450 $\mu\text{g mL}^{-1}$.
12 HEK293T was chosen to assess the potential application of this
13 new material in wound-healing, while HeLa was selected
14 because it is a widely used cell line for which a significant
15 number of cytotoxicity studies can be found in the literature.
16 No significant toxicity was observed, even at high concen-
17 trations of the solid (Figure S9). Only in the case of HEK293T,
18 a very low toxicity was found after 24 hours, and after 72
19 hours a decrease in the viability of the cells ($71 \pm 8\%$ at
20 450 $\mu\text{g mL}^{-1}$) was observed. However, as shown below, NO
21 loaded MIP-177 will produce a positive effect on cell
22 migration at concentrations of 90 $\mu\text{g mL}^{-1}$, for which no
23 toxicity was found. Compared with previous work, where all
24 porous materials studied for the same purpose presented at
25 least 20% toxicity with HeLa cells after 72 hours,^[33–35,43] no
26 toxicity was observed in the present case with HeLa cells (a
27 comparison with other MOFs is presented in the Supporting
28 Information).

29 MIP-177 stability, another critical requirement,^[45] was
30 evaluated under biological conditions. The Ti^{4+} concentration
31 in the liquid medium after 72 hours was 0.5 and 1.1 $\mu\text{g mL}^{-1}$ in
32 HeLa and HEK293T cell media, respectively, indicating a merely
33 residual degradation of the material (about 9% in the worst
34 case). The low degradation was confirmed by PXRD (Fig-
35 ure S12a) and by SEM images (Figure S12b). This is in
36 contrast with the MOFs tested so far, which have not been
37 stable enough in biological media (e.g., iron carboxylates
38 MIL-88's, MIL-100 and MIL-127 or metal(II) hydroxycar-
39 boxylate CPO-27(Mg, Zn) and/or MOFs containing very
40 toxic metals, such as Cu carboxylates or CPO-27(Cu, Mn, Ni
41 or Co)), hindering potential future therapeutic applica-
42 tions,^[12,14,32] particularly topical NO delivery to wounds. For
43 instance, biocompatible iron MOFs (Fe-MIL-100 and Fe-
44 MIL-127)^[32] were submitted for the same stability tests and
45 showed high concentrations of iron in both media after
46 72 hours (between 1.39 to 7.31 $\mu\text{g mL}^{-1}$, Table S1). This
47 instability may affect the controlled release of NO in
48 biological media (a spike in released NO is observed in the
49 first minutes upon contact with moisture, which renders those
50 MOFs unsuitable for medical applications).^[32] These results
51 underline the exceptionally low toxicity and excellent stability
52 of MIP-177 which are highly attractive qualities for topical
53 NO delivery applications.

54 To prove the efficiency of this new material as a viable
55 vehicle for the delivery of NO in biological applications,
56 a relationship between the amount of NO released and the
57 effect observed in biological systems must be clearly demon-
58 strated. So far, it has only been demonstrated in the literature
59 that NO released from porous materials can trigger some



positive responses in artery relaxation tests^[12,46] and in the inhibition of platelet aggregation, but the effect of the material concentration on the response has not been reported yet.^[7] Recently a NO-releasing zeolite was studied for wound healing. However, it showed a reduced effect, which is associated with the high toxicity of the NO-loaded material, likely due to its fast NO release (see Supporting Information for further discussion).^[18] In this case, we have chosen two relevant cellular processes that can be regulated by NO^[47,48] to demonstrate our ability to control them with NO-loaded MIP-177: (1) mitochondrial respiration, which was evaluated by measuring the oxygen consumption of cells, and (2) the wound healing process simulated using a cell migration assay.

Oxygen consumption was monitored in a respiration buffer before and after NO-loaded MIP-177 was added ($450 \mu\text{g mL}^{-1}$). A reversible inhibition of mitochondrial respiration by the released NO was observed (Figure 4b), reaching the highest inhibition rate of 78.5%, after which the respiration recovered to 70% of the initial value after 55 minutes. Results from NO and O₂ sensors (Figure 4a and b) were in agreement, showing that oxygen consumption was inhibited when the material was introduced and NO release

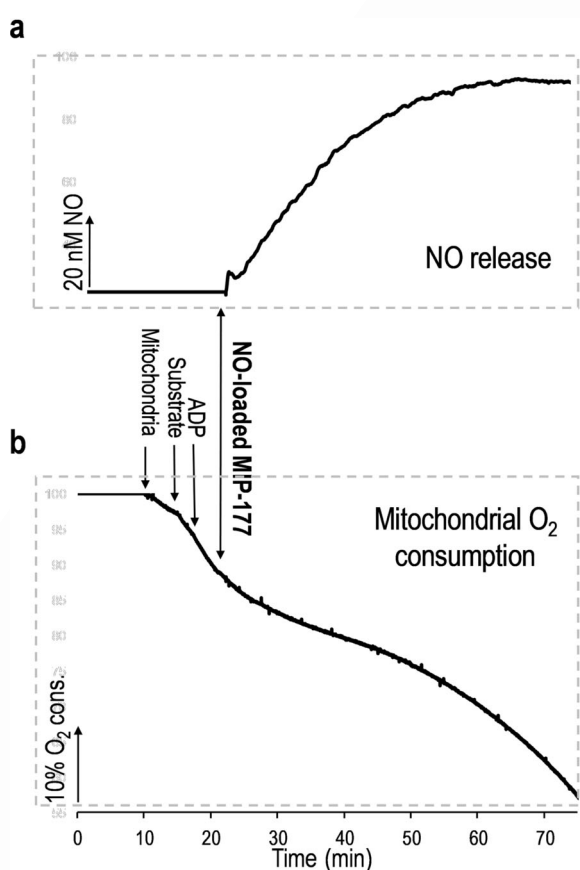


Figure 4. Effects on mitochondrial respiration in the presence of NO-loaded MIP-177. a) Release profile obtained from the NO-sensitive electrode using a MIP-177 concentration of $450 \mu\text{g mL}^{-1}$ in supplemented culture medium (RPMI-1640); b) Mitochondrial respiration assay using digitonin-permeabilized HeLa cells (1.25 mg of protein) with $450 \mu\text{g mL}^{-1}$ of NO-loaded MIP-177, tested by measuring the O₂ consumption with a O₂-sensitive electrode.

was started, and increasing further (Figure 4b) when the rate of NO release started to slow down (Figure 4a). At lower concentrations (Figure S13 and S14), the released NO causes a lower inhibition of the mitochondrial respiration rate ($\approx 18\%$ at $180 \mu\text{g mL}^{-1}$ and no inhibition at $90 \mu\text{g mL}^{-1}$). The high dependence of the mitochondrial respiration on NO concentration is demonstrated, proving that the NO released from MIP-177 indeed influences oxygen consumption and that it can be controlled by varying the concentration of material. These results are in agreement with the literature: low nanomolar concentrations of NO specifically inhibit cytochrome oxidase (the terminal complex of the mitochondrial respiratory chain) in competition with oxygen, and this inhibition is reversed when NO is removed.^[47,49]

It is reasonable to assume that this new NO donor could promote cell migration and wound closure and be a suitable therapeutic alternative in skin reparatory processes. To confirm such a possibility, the cell response to exposure with different concentrations of NO-loaded MIP-177 was evaluated using migration assays. After treating the cells at a concentration of $90 \mu\text{g mL}^{-1}$ (Figure 5a), an initial increase in the migration rate was observed ($5.5 \pm 1.3\%$) after 6 hours, achieving an improvement in wound closure of $8.4 \pm 1.4\%$ after 24 hours. These results suggest that a sustained effect could be achieved by renewal of the supply of NO-loaded MIP-177 to the medium every day, which is convenient for therapeutic purposes. Microscopic images (Figure 5), which show pre- and post-migration period cultures exposed to NO-loaded and unloaded (control) MIP-177, confirm the increased rate of wound closure with increased released NO. Some MIP-177 particle aggregates can be seen on the images, scattered outside the cells. Additional images that elucidate the particles in the cell cultures are shown in Figure S15). At higher concentrations of NO-loaded MIP-177 ($180 \mu\text{g mL}^{-1}$, Figure S16), no significant difference in the migration rate was observed. Since NO effects are extremely concentration dependent, perhaps the NO amount released from MIP-177 at $180 \mu\text{g mL}^{-1}$ was not suitable for this application, possibly due to an excessive supply of NO to the media (about the double amount of NO compared to $90 \mu\text{g mL}^{-1}$ sample shown in Figure S8). To the best of our knowledge, this is the first time that NO released from MOFs has been studied as an alternative for skin reparatory therapy. So far, only homogeneous NO donors have been considered for this therapeutic application. Zhan et al.^[50] reported that migration of HaCaT cells may be promoted either by Sodium nitroprusside (SNP), spermine NONOate, or *S*-nitroso-*N*-acetylpenicillamine (SNAP) as homogeneous donor in a dose-dependent manner. $10 \mu\text{M}$ SNP, $10 \mu\text{M}$ spermine NONOate and $100 \mu\text{M}$ SNAP were found to be the optimal concentrations to induce cell migration, promoting an increase of ca. 1–10% in the first 6 hours, in the same range as found here with MIP-177. Higher amounts presented inhibitory and cytotoxic effects were observed for the highest concentration ($1000 \mu\text{M}$). Other studies have also reported that exogenous NO supply using NO donors induces enhanced cell proliferation and migration in a concentration-dependent manner.^[51–53] However, current homogeneous donors can only provide non-targeted NO release since they

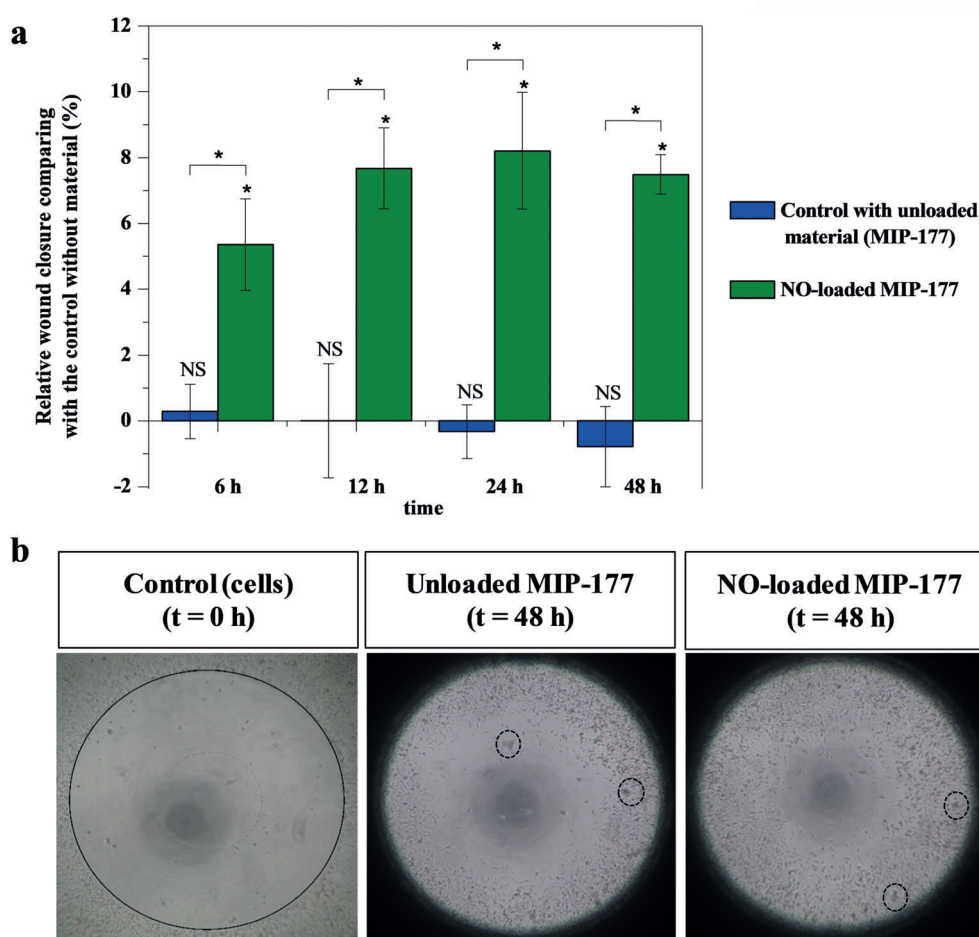


Figure 5. Effects of released NO on cell migration and wound closure. a) Cell migration data obtained by treating HeLa cells with MIP-177 as the NO donor material at a concentration of $90 \mu\text{g mL}^{-1}$. b) Microscope images captured during the cell migration assay. The first image on the left, captured without a detection mask in place shows the cells (control) in the pre-migration period (0 hours). The middle picture, captured with a detection mask, shows the migration of the control cells treated with unloaded MIP-177 post-migration (48 hours). The first image on the right, captured with a detection mask, shows the cells treated with NO-loaded MIP-177 post-migration (48 hours) ($90 \mu\text{g mL}^{-1}$). As shown in the images of cells post-migration, the presence of the scattered MIP-177 particles between the cells is evident (some examples outlined in black circles). Data in a) represents the average percentage migration \pm SD from three independent experiments, each performed in quadruplicate. Both conditions (cells treated with unloaded and NO-loaded material) are relative to the control (cells without material), that is, the difference in the percentage of wound closure between the tested condition and the control. Data were compared using the unpaired student's t-test to assess significance. * $P < 0.05$. Additional details in the Supporting Information.

can disseminate throughout tissues and body, which may affect their therapeutic efficiency and cause harsh side effects in vivo.^[6,54] Moreover, the decomposition products formed by the donors after releasing NO could be toxic and affect the surrounding tissues.^[55] In the case of MIP-177, a controlled release of pure NO is guaranteed, which, combined with its excellent stability and biocompatibility, are novel characteristics that support the use of MIP-177 as a suitable alternative in wound healing applications.

To satisfy specific medical requirements, the ability to adjust NO release kinetics and amounts of NO released from donors is highly desirable. In nanoporous carriers, these are dependent on pore structure and surface chemistry of the NO donor.^[6] The properties can vary drastically from one porous material to another. Thus, fine tuning of the same MOF structure would be an interesting approach to achieve specific modifications. To explore this concept, the same structure was

prepared, but with a partial isomorphous substitution of the Ti^{IV} metal centres with Fe^{III} (15% of Ti replaced by Fe), giving a new material with the same structure and two types of metal centres. Indeed, this preliminary study showed that this led to a slight increase of the stored amount of NO and, most importantly, that this approach could be further exploited to fine tune NO release capacity and kinetics (Figure S17).

Conclusion

The robust titanium carboxylate MOF MIP-177 features a new mechanism for NO adsorption/release through the formation of complexed nitrites on its inorganic building unit. MIP-177 furthermore shows a combination of excellent biocompatibility, particularly at therapeutically relevant concentrations, and unprecedented stability in biological media,



which has led to a remarkably slow NO release in biological fluids. This facilitates control over biological responses that exhibit a dependence on the concentration of NO. These conclusions provide a step forward in the development of advanced nanoporous materials for NO delivery for use in wound healing therapy. MIP-177 can indeed be considered as a new component in the development of NO delivery platforms, to be potentially deployed in combination with polymers, gels and ointments for future practical clinical applications.

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Conflict of interest

The authors declare no conflict of interest.

Keywords: IR spectroscopy · metal–organic frameworks · modelling · nitric oxide · wound healing

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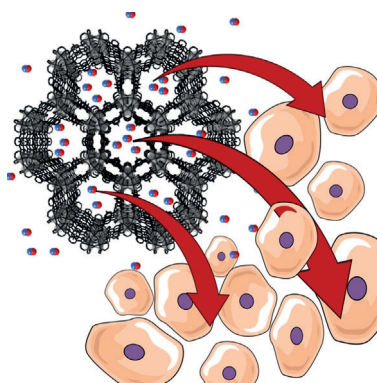
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Research Articles

NO Release






R. V. Pinto, S. Wang, S. R. Tavares, J. Pires,
F. Antunes, A. Vimont, G. Clet, M. Daturi,
G. Maurin, C. Serre,*
M. L. Pinto*    

Tuning Cellular Biological Functions
Through the Controlled Release of NO
from a Porous Ti-MOF

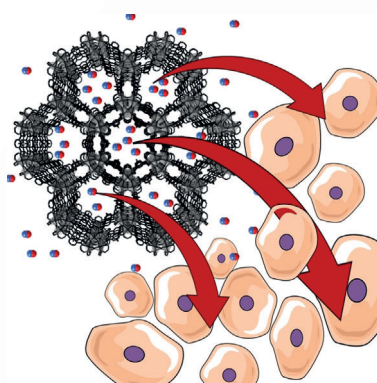


NO limit: The delivery of nitric oxide with a new titanium metal–organic framework allows the control of biological functions of cells, due to its adsorption/release mechanism based on the formation of nitrites bound to the titanium sites. The high storage capacity, biocompatibility and slow release and stability in biological media, render this material an exceptional vehicle for nitric oxide delivery.

NO-Freisetzung

R. V. Pinto, S. Wang, S. R. Tavares, J. Pires,
F. Antunes, A. Vimont, G. Clet, M. Daturi,
G. Maurin, C. Serre,*
M. L. Pinto*     

Tuning Cellular Biological Functions
Through the Controlled Release of NO
from a Porous Ti-MOF



Die Freisetzung von NO durch ein Ti-metallorganisches Gerüst ermöglicht aufgrund des NO-Speichermechanismus die Kontrolle der biologischen Funktionen von Zellen. Der Mechanismus beruht auf der Bildung von Nitriten, die an die Titanzentren gebunden sind. Die hohe Speicherkapazität, Biokompatibilität und die langsame Freisetzung und Stabilität in biologischen Medien machen dieses Material zu einem interessanten Werkzeug zur Abgabe von NO.



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