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Review of tellurite glasses purification issues for mid-IR optical fiber applications

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Abstract: We reviewed the tellurite glass purification studies performed over the last 20 years, in particular dealing with the mid-infrared transmission performances. Best results are obtained under rich O₂ atmosphere with platinum crucibles although Gibbs energy calculations show that other crucibles material are possible such as gold or alumina. -From the point of view of the thermodynamics of involved oxidation reactions only, we identified the suitable conditions based on Gibbs energy calculations to be a synthesis temperature set above 900K and a O₂ pressure around 10⁵ Pa. The performances of the different purification technics were also compared. Finally, we analyzed the tellurite fibers recently optimized for mid-infrared which present low attenuation up to 4 μm allowing supercontinuum generation up to 5 μm.

1 Introduction

Tellurite glasses have been investigated for more than 150 years with first observations of the potential glass forming ability of tellurium oxide reported as early as 1834.^{1, 2} The

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publication of Stanworth³ was the first one fully dedicated to tellurite glasses. Since then, many different TeO₂-based glass systems have been studied and the first developments of tellurite fibers were reported in the 1990s.⁴

Tellurite glass fibers are of strong interest over a wide range of infrared (IR) wavelengths as waveguides or gain media for fiber lasers and fiber optical amplifiers based either on rare earth doping⁴⁻⁶ or Raman gain.^{7, 8} In addition, tellurite glasses can be used as broadband laser sources through supercontinuum generation for medical, environmental, and defense applications.⁹⁻¹¹ These glasses offer many advantages compared to other materials for the IR spectral region. For example, chalcogenides based on S/Se/Te can transmit light up to 20 μm and beyond, depending on the composition.¹² These materials have high optical nonlinearity, which makes them good candidates for mid-infrared (mid-IR) applications. However, their composition generally includes highly toxic elements such as arsenic. In addition, chalcogenide glasses have low physical strength and glass transition temperatures (<200°C), together with poor resistance to moisture¹³⁻¹⁵, thus limiting possible applications. Halides and, in particular, fluoride glasses are another family of glasses useful in the Mid-IR thanks to their wide transmission window (up to 9μm). Fluoride glasses can also act as hosts for rare earth doping. However, fluorides are physically weak compared to silica fibers and have a low glass transition temperature (260°C). They also suffer from poor chemical durability, which results in very few glass compositions that can be handled at room atmosphere or processed into fibers. Thanks to their reasonably wide transmission window (0.4-6.5 μm), moderate glass transition temperature (<400°C), and high rare earth solubility, tellurite glasses offer an interesting compromise compared to fluorides and chalcogenides. Their range of possible chemical compositions facilitates optical design through refractive index and dispersion profile management. Moreover, tellurite glasses

have the widest range of nonlinear responses (Raman gain, $\chi^{(3)}$) among heavy oxide glasses with higher refractive indices and nonlinear coefficients

The need to minimize water contamination in tellurite glasses was obvious and significant reductions in OH impurity contents were made in the 2000s.¹⁶⁻²³ Tellurite glass purification, especially hydroxyl group removal, is of importance for many optical fiber applications. Because OH groups occupy a variety of different sites in TeO_4 and TeO_3 units, OH-related absorption blocks transmission in a large band from 2.7 μm up to 4 μm . Applications requiring transmission in the mid-IR, such as supercontinuum (SC) light sources, require a significant decrease of absorption in this range of wavelengths and extension of the transmission range to the multiphonon limit of TeO_2 -based materials (above 4.5 μm).²⁴
²⁵ As a consequence, only a few laboratories were able to demonstrate the spectral broadening up to this limit in the mid-IR.^{11, 26-28} As a side note, rare earth-doped glasses used for emission around 2 μm are also sensitive to OH content, which is responsible for the quenching of photoluminescence.²⁹ The glass structure is modified by OH impurities, which lead to non-bridging oxygens, similar to alkali-containing glasses, and result in a decrease in network formation as well as a decrease of the transformation range viscosity and glass transition temperature.²⁰ However, even under the worst conditions, i.e., when glass synthesis occurs in the ambient atmosphere, the OH content in glass remains low and in any case below 150 ppm by weight as discussed below.

Tellurite glasses can be purified either by physical dehydration through careful control of the synthesis conditions (e.g., atmosphere, drying of the precursors), or by the use of desiccating additives (e.g., halides such as NaCl, NaF, ZnF_2 or ZnCl_2). Another possibility to purify these glasses is to use reactive atmosphere processing (RAP) techniques. The purpose

of this paper is to review different tellurite glass purification techniques and summarize the state-of-the art absorption/attenuation results.

2 Composition

Glass transparency is intrinsically governed by composition; tellurite glasses have been investigated because of their wide transparency in the mid-IR range. Long-wavelength transmission such as mid-IR is limited by absorption related to chemical bond vibrations, which are also called phonon energy. In the case of TeO₂ based glass, the theoretical limit is set by TeO₂ itself. Hence, for shifting the absorption limit as far as possible towards long wavelengths, the glass composition must only include constituents (modifiers and intermediates) whose inherent absorption wavelength is longer than that of tellurium (IV) oxide, which is a strategy used by several groups.^{11, 30} Table 1 gathers fundamental absorptions and the three first harmonics of several oxides commonly found in glass compositions. Network modifiers such as alkalis, which promote formation of non-bridging oxygens, give rise to TeO₃ units that absorb at 740 cm⁻¹ compared to 650 cm⁻¹ for TeO₄ units. This shift results in a decrease in the mid-IR transmission window. Additionally, hydrogen can then associate with non-bridging oxygens to form hydroxyl groups, which give rise to unwanted absorption at the hydroxyl fundamental absorption peak around 3400 cm⁻¹.

3 Sources of Impurities

Optical losses are largely determined by: 1) the concentration of impurities in the starting materials; 2) impurities imparted during glass synthesis; and 3) scattering from inclusions such as crystalline phases and bubbles.^{7, 16, 32} Among the extrinsic impurities present in the starting products, hydroxyl groups and 3d-transition metals actively absorb in the infrared and visible ranges and are particularly detrimental for optical applications in mid-IR.

Tellurite glass preforms for optical fibers are usually manufactured by melt-casting from open platinum or gold crucibles. As such, the ambient atmosphere is the main source of OH impurities. Oxide precursors are another source of OH⁻ contamination. Their metallic purity, mainly transition metals, is specified by the manufacturer, but other impurities such as water or carbon are not specified. The water content of the precursors depends on their synthesis process (extraction and refining). For example, carbonates are an important source of moisture. Contrary to the synthesis atmosphere for which we can control the water content the raw materials OH content is unknown and difficult to remove especially if it is chemisorbed.^{33, 17} As a matter of facts, several study used homemade precursor to have a better control of their impurities content in order to improve optical properties of tellurite glass.^{7, 17, 16}

4 Determination of OH concentration

Hydroxyl groups strongly absorb in the mid-IR, between 2.7 and 4 μm. The extinction coefficient (ϵ) varies from 34 to 90 L•mol⁻¹•cm⁻¹ for different studies³⁴. At least in part, these discrepancies are related to calculating the extinction coefficient from either an exponential Beer Lambert law or an absorbance law. For example,³⁵ absorbance measurements in silica glass resulted in an OH extinction coefficient of 42 L•mol⁻¹•cm⁻¹ at 2.8 μm, in accordance with data published by other authors.^{34, 35} In the present work, absorption coefficients were determined using the exponential Beer Lambert law.

Because OH extinction coefficients have not been widely measured for tellurite glasses, values often come from measurements in silica glasses. However, data are available for tungstate tellurite compositions with an average value of $\epsilon(\text{OH}) = 109 \text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ (from exponential Beer Lambert law) around 3000 cm⁻¹ (3.33 μm).^{36, 37} The OH extinction

coefficient at 3 μm in oxide glasses is often assumed to be $113 \text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ ³⁸ which fits well with the one measured in tungstate tellurite glasses.^{36, 37} We calculated the residual OH concentration [OH] (in wt ppm) using the extinction coefficient measured for tellurite tungstate glasses ($109 \text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$) and the absorption coefficient measured around 3 μm α_{OH} in cm^{-1} into Equation 1:

$$[\text{OH}] = \frac{\alpha_{\text{OH}} \times M_{\text{OH}}}{10^{-3} \times \rho \times \varepsilon(\text{OH})} \quad (1)$$

where M_{OH} is the molecular weight in $\text{g}\cdot\text{mol}^{-1}$ of OH and ρ is the volumetric mass density of tellurite glass in $\text{g}\cdot\text{cm}^{-3}$.

Removal of OH groups represents the most important challenge to maximize the transmission window of tellurite glasses and take full advantage of materials in bulk or as fibers. The most common approaches for purification of tellurite glasses are removing water and hydroxyl groups present in the precursors either before or during synthesis.

5 Crucible and redox conditions of the synthesis atmosphere

Different kinds of crucibles can be used for tellurite glasses synthesis. The common types are platinum (Pt), gold (Au) or corundum (Al_2O_3). Using the Ellingham approximation, the standard Gibbs' free energy ($\Delta_r G^0$) for oxidation of these metals (Pt, Au and Al) as a function of temperature was compared to the one for Te, see figure 1(a). The standard enthalpy and entropy of reaction were assumed to be independent of the temperature in ranges where no phase change occurred (see table 2).

At the phase change temperatures, data for fusion enthalpies and entropies were taken from.³⁹⁻⁴¹ The calculations supported the use of platinum, gold or corundum crucibles. Tellurium oxidation is more favorable compared to Pt or Au oxidation, which means that TeO_2 will not be reduced when heated in Pt or Au crucibles. As shown by the Ellingham

diagram figure 1(a)), Al_2O_3 reduction is not favored, which indicated that corundum should be inert in contact with TeO_2 melts even in presence of Te traces.

By examining the evolution of the O_2 equilibrium pressure $P(\text{O}_2)_{\text{eq}}$, deduced from $\Delta_r G^0$, as a function of temperature for oxidation of Te to TeO_2 and Pt to Pt_3O_4 (figure 1(b)), suitable working oxygen activities can be estimated to ensure that TeO_2 does not reduce to Te while simultaneously ensuring that the Pt crucible does not oxidize. Reduction of TeO_2 is detrimental because it leads to an increase in background absorption.^{7, 8} According to figure 1b, assuming a synthesis temperature of 700°C , O_2 pressures up to about 10^5 Pa (1 bar), more specifically in the range of $9.840 \times 10^4 < P_{\text{O}_2} < 1.006 \times 10^5$ ($10^{-0.007} < (P_{\text{O}_2}/P^\circ) < 10^{0.0025}$), ensures that Pt does not oxidize and TeO_2 does not reduce. For example, an atmosphere of pure oxygen at ambient pressure would avoid reduction of TeO_2 .^{7, 8} Pressures of O_2 higher than about 10^5 Pa should be avoided to prevent Pt oxidation. Since nitrogen does not react with tellurite glasses, dry air is another possible processing atmosphere. Indeed, for a total dry air pressure of 5×10^5 Pa, P_{O_2} is within the foregoing defined window.

Platinum is not totally inert in contact with TeO_2 , as mentioned by Churbanov et al.¹⁶
¹⁷ Platinum has an effect on glass color due to a slight reaction of the crucible material under synthesis conditions, see figure 2. For the same glass composition, synthesis in platinum crucible leads to a slightly yellow colored glass while a synthesis in gold crucible results in transparent glass. The color may be related to dissolved Pt, but does not appear to impact the IR range^{16, 17}. According to a study of phosphate-based glass⁴², Pt dissolves as Pt^{4+} under oxidizing conditions. This has also been observed for paratellurite (TeO_2) crystals⁴³ synthesis, where it was assumed that Pt was oxidized and then dissolved as Pt^{x+} while some TeO_2 was reduced to TeO, which produces black inclusions. During the growth of these TeO_2 crystals, Pt_xTe_y alloys may also be formed.⁴³ Single crystal growth experiments were performed under

air, vacuum, or argon whereas TeO₂ reduction can be avoided by working under an O₂ rich atmosphere. The present calculations emphasize the need for an O₂ rich atmosphere during tellurite glass synthesis. However, the mechanism of Pt interactions during glass synthesis still needs to be clarified. Similarly, gold may also dissolve into the melt by corrosion of the crucible at the synthesis temperature of 700°C.^{44, 45}

6 OH-groups content control

One way to remove OH groups present in the starting materials is to heat treatment them. The precursors can be placed under vacuum or under a flow of dry oxygen to promote dehydration.^{17, 46-50} However, figure 1b shows that low O₂ pressures, such as those produced in vacuum, result in TeO₂ reduction. Thus, redox conditions should be considered when selecting drying conditions; obviously, O₂ pressures far below 10⁵ Pa (1 bar) will lead to some TeO₂ reduction, which will increase background absorption. Nevertheless, Mori and al^{7, 8} showed that drying of starting materials can decrease losses in the near IR in tellurite glasses.

Another way to remove water and OH groups from tellurite glasses is to treat the liquid bath with reactive species during synthesis^{29, 46, 47, 51}. The usual process for the purification of tellurite glasses is to use desiccant halides reagents. Fluorinated or chlorinated forms of metals such as ZnF₂^{19-21, 50, 52-57}, ZnCl₂^{58, 59}, NaF⁵⁵, NaCl⁶⁰, PbCl₂^{50, 60} or BaF₂, PbF₂ and CaF₂⁵⁰ can be mixed with oxide precursors to remove OH groups from tellurite glasses. These halide species react with free water or OH groups during synthesis to form HCl or HF gases, which then leave the liquid bath. A less common process is to carry out a chlorination/fluorination treatment of the starting powders before the synthesis. This can be done through the use of a solid reagent (NH₄F-HF) or through reactive atmosphere

processing (RAP) with a reactive gas, as detailed further in the section 7.3. These various purification techniques have extended light transmission in tellurite fibers above 4 μm , beyond the water-related absorptions.

7 Purification processes

Characteristic absorption bands found in tellurite glasses are listed in table 3, including extrinsic absorption due to OH impurities and intrinsic absorption that limits glass transmission below 6.5 μm . The processes described below are performed by means of glove boxes or dedicated reactors.

7.1 Controlled atmosphere

Fusion of tellurite precursors must be done under an oxidizing and dry atmosphere.^{21, 33} An atmosphere rich in O_2 , but dry (water concentration less than 0.5 ppm vol) enables reduction of the OH concentration by a factor of thirty (from 150 ppm to 5 ppm) as compared to synthesis in open air, as shown in figure 3.^{21, 33} This reduction in OH concentration pushes the fiber transmission limit from 1.8 to 2.8 μm , as shown in table 4, which reports methods implying only a controlled atmosphere (no reagents, no RAP and no homemade precursors) and figure 9. However, this synthesis route is not sufficient to reach the intrinsic limit of tellurite glass transmission that is controlled by multiphonon absorption resonances beyond 4.3 μm (see Table 3). Indeed, an absorption coefficient of 0.25 cm^{-1} (calculated from Beer Lambert law, figure 3) corresponds to a loss level of about 108 dB/m, which prevents transmission of any signal above 3 μm for a meter-long fiber segment.

7.2 Homemade chemicals

Mori et al^{7, 8} developed tellurite glass fibers for telecom applications. To decrease optical losses in the telecom bands, they used homemade ultrapure TeO_2 . As a result, they

obtained ultra-low loss tellurite fibers with an attenuation below 27 dB/km at 1.23 μm , while it was about 900 dB/km with commercial TeO_2 . By adding a vacuum degassing step of raw materials, the same loss level was reached at 1.55 μm . In this case, the TeO_2 was fabricated from metallic Te using a wet chemical process.^{7, 68}

The Institute of Chemistry of High-Purity Substances has developed its own techniques, based on oxide distillation, to prepare TeO_2 as well as other oxides.^{16, 17, 32, 36, 49, 69, 70} As reported in table 5, using such purified precursors decreased OH content and led to an absorption coefficient below 0.001cm^{-1} between 2.7 and 4 μm , which means that the residual OH impurity concentration was below 0.1 ppm wt.

Figure 4a gathers obtained bulk absorption coefficients for different studies that synthesized starting materials. As shown in figure 4b, tellurite glass preparation from ultrapure precursors is attractive for telecom applications. Note that Mori et al⁷ also reported tellurite-based fiber Raman amplifiers as well as erbium doped fiber amplifiers. By synthesizing starting materials, bulk absorption coefficient corresponding to OH absorption as low as 0.001 cm^{-1} can be reached, which means that the corresponding fiber could exhibit losses that would not exceed 0.5 dB/m between 2.7 and 4.3 μm . However, measurement of fiber attenuation in this wavelength range is not available for several reasons. First, bulk measurements are made on millimeter thick specimens, so the uncertainty increases when extrapolating the results to meter or kilometer lengths. Also, the thickness of the bulk is not sufficiently large to give an accurate loss value when the OH loss is reduced while possible contamination from polishing media may also affect measurements.^{60, 16} Secondly, several fabrication steps are needed to obtain fiber. Fiber drawing, for example, requires heating of the glass preform above the glass transition temperature. Depending on process conditions

(furnace atmosphere, time and rate of heating) and the glass composition, this step can be responsible for increasing losses.

As we show later in this paper, very few studies have reported fiber (i.e. at least meter-long specimens) loss measurements above 2.5 μm and all of those results have been obtained through glasses purified by the use of desiccant reagents.

7.3 Reactive atmosphere processing (RAP).

Gaseous reagents can be added to the synthesis atmosphere during fusion or a separate pre-treatment step to eliminate the water-related impurities from the raw materials. This method is rarely used because of the complexity of handling halogens like F_2 or Cl_2 . For tellurite glasses, the majority of the reported studies have used chlorine.^{29, 46, 47, 51} Another drawback of this method is unwanted halidation reactions that can occur in the melt in addition to the dehydration reactions. For instance, chlorine can also react to form chlorides resulting in an oxyhalide glass.

Ideally, chlorine will react with hydroxyls present in compounds such as $\text{TeO}_2(\text{OH})_2$, $\text{Zn}(\text{OH})_2$ and $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$.⁴⁶ Ellingham diagrams were plotted considering the reaction of 1 mole of Cl_2 with several hydrated precursors in the TeO_2 - ZnO - Na_2O (TZN) glass system, see figure 5. Thermodynamic data such as standard enthalpy of formation, standard entropy and standard heat of fusion were taken from published sources³⁹⁻⁴¹ and hydrated precursors were considered when data were available. However, temperatures of phase transition or decomposition were not always available, so the validity of changes in Gibbs' energy at the higher temperature may be questionable. Results for $\text{Zn}(\text{OH})_2$ and $\text{Na}_2\text{CO}_3 \cdot (\text{H}_2\text{O})$ are plotted with dotted lines above their decomposition temperatures of 125 and 100°C, respectively.³⁹ These curves give information about the synthesis conditions that favor dehydration. According to the Ellingham diagrams (figure 5), $\Delta_r G^\circ$ is negative for all of the dehydration

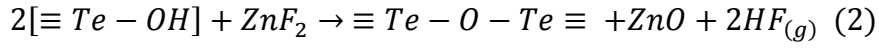
reactions with chlorides above 637K, indicating that dehydration is favorable above this temperature. The formation of NaCl and ZnCl₂ are favorable at all temperatures.

Limiting the formation of oxyhalide species in the glass requires limiting the duration of the chlorine gas treatment, with no guarantee that chlorine will react exclusively with water or OH groups. The same considerations are valid for fluorine (F₂, not shown here), which can form oxyfluoride glasses. Halide and oxyhalide glasses are more sensitive to the ambient atmosphere, especially moisture, which makes these materials difficult to process, particularly for fiber drawing⁷².

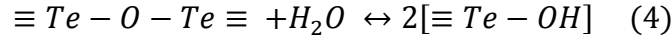
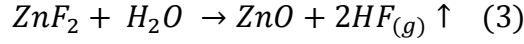
Figure 6 shows the absorption coefficient obtained for glasses using RAP; the lowest OH absorption coefficient was 0.05 cm⁻¹ obtained by Joshi et al⁴⁶ corresponding to 21.7 dB/m. However, to the best of our knowledge, no reports of fiber produced from glass prepared through RAP have been published. Moreover, the value of 21.7 dB/m is still too high to consider meter-long fiber applications at 3 μm and above.

7.4 Desiccant reagents

The addition of solid halogen-containing compounds is preferred to avoid using toxic and reactive halogen gases in the processing of tellurite glasses.⁶⁰ This technique consists of using a reagent which will react with free water or OH groups in the starting chemicals and glass batch during synthesis. Halide dehydrating reagents that have been used include ZnF₂, PbCl₂ and TeCl₄. As shown in table 7, this is the most common purifying process for tellurite glasses and the most used halide is ZnF₂, which reduces OH content by Reaction (2).¹⁹ Fluorides are known to be highly reactive with water and OH groups.⁷³ One of the reaction products is zinc oxide, which is part of several tellurite glass compositions.¹⁹ $2[\equiv Te - OH] + ZnF_2 \rightarrow \equiv Te - O - Te \equiv + ZnO + 2HF_{(g)}$ (



In addition, halide reagents can reduce the water in the glass melting atmosphere (Eq (3)), which prevents hydrolysis of the melt (eq (4)).



Reactions with halides can also impact the crucible choice. To address that, we have plotted an additional Ellingham-like diagram using the same methodology as described above. In this case, the reaction of 1 mole of ZnF_2 with four hydrated precursors of the TZN glass system were considered, see figure 7(a).³⁹⁻⁴¹ TeO was included in this analysis as it may be present in the starting powder and increases background attenuation⁷. Hydrated precursors for which thermodynamic data were available were analyzed, but heats of fusion and temperatures of phase transitions or decomposition were not available for these compounds. These predictions provide information about the synthesis conditions that favor dehydration. The curve relative to sodium carbonate decahydrate ($Na_2CO_3 \cdot 10H_2O$) is not presented here contrary to the work of Joshi et al⁴⁶ because it decomposes at 34°C according to.³⁹ The curve for sodium carbonate monohydrate ($Na_2CO_3 \cdot H_2O$) is plotted with a dashed line above 100°C, which is its reported decomposition temperature.³⁹ As a result of decompositions and gaps in data, the thermodynamic considerations presented here are notional and not intended to be quantitative. It is worth noting that decomposition of hydrated carbonates produces water molecules that can then react with other compounds, especially TeO_2 (see figure 7b). Our goal here is to provide information about thermal/gas atmosphere conditions, which allow for dehydration of the starting materials used for tellurite glass synthesis. To that end, $\Delta_r G^\circ$ was calculated from room temperature to 850°C,

which corresponds to the range of temperatures used for tellurite glass synthesis. According to figure 7(a) the four dehydrating reactions are favored at higher temperatures. At the usual synthesis temperature (around 850°C, see table 7), $\Delta_r G^\circ$ is significantly negative ($\Delta_r G^\circ < -150$ kJ/mole), which indicates that dehydration is favored for all the four hydrated starting products. Moreover, according to figure 7b, above 563 K (290°C) the hydration of TeO_2 is not favorable. The same figure shows that Reaction (3) is favored above 823 K (550°C), which means that ZnF_2 is an effective additive for removing water. Hence, synthesis needs to occur above 550°C, and then glasses need to be quenched through the zone where dehydration is thermodynamically favored and kinetically significant to maximize the dehydration processes. Figures 7(c) and (d) show the log of the HF pressure and the equilibrium constant as functions of temperature. The equilibrium constant depends mainly on the gas pressures when the activities of the solid and pure liquid phases are assumed to be 1.

Figures 7(a) and 7(c) show that high temperature and moderate HF pressure favor the reaction with ZnF_2 . This means that evacuating the HF as it is formed, which is the case when glass is prepared under flowing gas, will shift the equilibrium toward the dehydration. This may be problematic when working under gas pressure in static conditions.⁸³ In this case, a large reactor volume or regular refreshing of the gas is required to maintain a low partial pressure of HF. In the case of the reaction with $\text{TeO}_2(\text{OH})_2$, interpretation of the equilibrium constant evolution is not as apparent. Experimentally, only the O_2 pressure can be controlled. In many studies, pure O_2 was used, which means that the O_2 pressure is slightly above 10^5 Pa (1 bar). As in flow conditions, HF is evacuated as it is formed and its pressure should stay as low as possible to allow for O_2 pressure to stay above 1 bar. To estimate the O_2 pressure threshold, we consider the worst case, i.e. the formation of all the HF at the same time, which means an instantaneous reaction of all the desiccant reagents. The

experimental conditions used in our glass preform fabrication can be approximated using 5 at% ZnF₂ in 40 g of TZN glass (80TeO₂-10ZnO-10Na₂O, % at) in a volume of 3L. For this estimation, activity of non-gaseous species was assumed to be 1 and gases were assumed to be ideal. At 973K (700°C), the HF pressure would be 75000 Pa (0.75 bars) leading to a threshold O₂ pressure of 315000 Pa (3.15 bars, calculated from the equilibrium constant shown in figure 7(d)), which means that above this threshold the reactions that remove hydroxyls are no longer favored. As mentioned above, working under flow conditions, allows for HF removal as it is formed and guarantees that the O₂ pressure is well below the calculated threshold as long as all of the ZnF₂ does not react instantaneously.

Based on previous reports and our own work^{83, 84}, fluorides (especially ZnF₂) seem to be more effective than other halides, see figure 8. Feng and Liao^{60, 75} used up to 20 at% halides to reduce OH absorption, while Savelii, Strutynski, or Rhonehouse^{21, 28, 30} used no more than 5 at% of ZnF₂. According to figure 8(a) and 8(c), lower amounts of reagents (e.g., <5 at%) are sufficient to maximize the decrease in OH-absorption. Moreover, with an optimized process, up to 5 at% of ZnF₂ does not degrade glass properties as suggested by³⁰ and shown by the refractive index and T_g measurements of Strutynski et al²⁸, even if some residual halide is present in these glasses. A non-zero residual amount of fluoride indicates that substantially all of the OH groups were removed. In this case, this small residual amount of fluoride does not degrade the chemical durability of the glass nor does it lead to breakage of tellurium oxygen bonds that would weaken the glass network and increase absorption in the mid-IR. Higher initial amounts of halide can show better results on bulk sample as reported by Jia^{10, 78}, see figure 8(c). However, glasses with substantial halide content are no longer tellurites, but fluorotellurites, since the only oxide in the composition is the glass former (TeO₂) while the other constituents are fluorides.

The addition of fluorides leads to composition modifications as shown by decreases in refractive index and T_g along with changes in Raman spectra.^{19, 21, 30, 52} Fluoride additions can also prevent or complicate post processing of the glass including fiber drawing. Any excess halide becomes part of the glass network and halide glasses, especially fluorides, are very difficult to draw.⁷² Additions of 5 and 10 at% ZnF_2 to a TZN glass revealed that the composition containing 10% ZnF_2 is much more difficult to process and yields fibers with losses that were too high to be measured by the cut back technique using specimens of ~20 cm long.^{83, 84} Very few studies have reported fiber drawing and attenuation measurements on tellurite glass fibers prepared through reagent additions, especially for wavelengths above 3 μm . Figure 9 summarizes reported attenuations for these tellurite fibers made from a glass prepared through the desiccant reagent technic.

Recently Yao *et al.* prepared a step index fiber based on fluorotellurite glasses (for core and cladding) that exhibited low losses between 1 and 5 μm ⁸² as shown by the blue curve in figure 9a. However, IR transmission of fluorotellurite glasses is limited by TeO_2 phonon resonances, which occurred at shorter wavelengths than fluoride alone, while the non-linearity was reduced compared to tellurite glasses. Most fluorotellurite glasses suffer from poor stability against crystallization and water attack, leading to delicate glass processing, especially the fiber drawing. However, for some compositions, the mechanical and thermal stability of fluorotellurites can be better than for fluoride and chalcogenide glasses. In addition, fluorotellurites can also have a T_g above 400°C, which is higher than usual fluoride, chalcogenide and tellurite glasses.⁸² When using fluorides as desiccant reagents, the goal is for the final glass to not contain any fluoride.³⁰ Thus, the final glass is expected to be a pure oxide one, in the present case a tellurite glass.

Considering OH reduction and infrared signal transmission above 3 μm , the best attenuation has been reported by Strutynski *et al.* in step index tellurite fibers^{27, 28} (orange curve in figure 9(a & b)). The core of this fiber was a TZN glass with 5 at% of ZnO replaced by ZnF₂. Attenuation was less than 1dB/m from the visible up to 3 μm , and the attenuation stayed below 8 dB/m up to 4 μm . Absorption above 4 μm is due to an intrinsic limitation based on TeO₂ multiphonon absorption. Note that when authors have not specified whether their measurements were made on bulk glass or fibers, data were not plotted in figure 9, such as for reference.¹¹ Based on such optimized fiber transmissions above 3 μm , enhanced SC generation in tellurite fibers was demonstrated by different teams, as depicted in figure 10.

More specifically, these tellurite fibers exhibited signal transmission up to 4 μm with loss levels below 10 dB/m from visible up to 4 μm . In each cases, ZnF₂ was used as desiccant reagent.^{27, 28, 30} Two SC spectra were generated with the same pumping laser source in two different fibers, which correspond to red and black plots in figure 10. These results obtained by two different collaborations underline the effectiveness of tellurite OH removal for mid-IR SC generation. Note that we recently obtained a similar SC in another tellurite glass fiber prepared through the same process, but with lower coupled peak power.⁹ The larger power spectral density is due to the ps pumping, this is why dynamic is larger (40dB) for the green curve than for the red and black ones (20 dB). Figure 10 also shows the SC generated in the tapered fluorotellurite⁸⁵ (blue curve) fiber whose losses are presented in blue in figure 9a. The addition of more than 20 at% fluoride reduced losses in the mid-IR spectral range (<2.3 dB/m from 1.1 and 4.5 μm). On the other hand, despite the use of a much higher peak power pump (160 kW launched), the generated SC does not extend further into the IR, which may be due to the lower nonlinear absorption properties of oxyfluorides compared to

pure tellurite glasses. In our case, these results were obtained with alkali containing glasses. However, as discussed above, alkalis have to be removed from glass to optimize mid-IR glass transmission because they promote formation of non-bridging oxygens that are responsible to the blue shift of phonon absorption and an increased sensitivity to hydroxyl-group absorption. Therefore, an alkali-free tellurite glass will further improve the fiber attenuation and the mid-IR SC extension.

Other processes have been used for glass synthesis including chemical vapor deposition (CVD) and microwave melting. CVD is complex and has primarily been used for silica glasses for the manufacture of telecommunication fibers. Microwave melting has been implemented in multicomponent glasses such as silicates⁸⁶ and chalcogenides⁸⁷. In both cases, microwave melting led to lower OH contents compared to furnace melted glasses without additional purification steps. According to⁸⁸, zinc oxide is a strong microwave absorber at 2.45 GHz above 600°C. Therefore, tellurite glasses containing ZnO can be synthesized by combining conventional and microwave heating. However, to the best of our knowledge, no results on tellurite glasses have been reported on microwave-assisted synthesis.

8 Summary and outlook

Gibbs free energy calculations in the framework of Ellingham approximations provide valuable information about synthesis conditions of tellurite glasses, as well as for dehydration reactions for water-containing raw materials. Synthesis temperatures above 900K together with O₂ pressures approaching 10⁵ Pa (1 bar) ensure oxides precursors, such as TeO₂, do not reduce and oxidation of platinum crucibles is avoided. Purification of tellurite glasses requires processing under a controlled atmosphere to favor dehydration and to

eliminate moisture from the synthesis atmosphere, and usually glove boxes or dedicated reactors are used. Higher purity extends the tellurite fiber transmission window to above 4 μm , to the intrinsic multiphonon absorption limit. One purification method is to use ultrahigh purity precursors to perform the synthesis, which results in bulk samples containing the lowest OH levels ever reported. However, no absorption data at 3 μm and above have been published on fiber resulting from this glass synthesis process. This could be due to the difficulty to prepare enough of highly pure precursor for a fiber glass preform combined with additional processing steps responsible for impurities in the resulting fibers. Indeed, such ultra-high purity raw materials are not commercially available and raw material purification is resource intensive.

This review of synthesis and purification processes reveals that the best combination of fiber drawing and mid-IR optical transmission for tellurite glasses are obtained when using a zinc difluoride as desiccant reagent, avoiding its excess not to alter the final glass composition, and ensuring removal of HF as it is formed. Such tellurite fibers exhibit an infrared supercontinuum extending from 1 μm up to 5 μm . This opens up new possibilities for applications of tellurite fibers such as absorption spectroscopy⁸⁹, frequency comb generation for high resolution spectroscopy⁹⁰⁻⁹², optical coherence tomography^{93, 94}, food quality analysis⁹⁵, biomedical endoscopy⁹⁶⁻⁹⁹ and atmospheric gas sensing.^{81, 100}

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9 List of tables

Table 1: Phonon energies of several oxides^{30, 31} and corresponding absorption wavelengths

	Phonon Energy (cm^{-1})	Fundamental Absorption (μm)	First harmonic (μm)	Second harmonic(μm)	Third harmonic (μm)
B₂O₃	1400	7.1	3.6	2.4	1.8
P₂O₅	1200	8.3	4.2	2.8	2.1
SiO₂	1100	9.1	4.5	3.0	2.3
WO₃	925	10.8	5.4	3.6	2.7
GeO₂	900	11.1	5.6	3.7	2.8
Nb₂O₅	860	11.6	5.8	3.9	2.9
TeO₂	650	15.4	7.7	5.1	3.8
Y₂O₃	590	16.9	8.5	5.6	4.2
Ta₂O₅	550	18.2	9.1	6.1	4.5
Ga₂O₃	500	20.0	10.0	6.7	5.0
La₂O₃	450	22.2	11.1	7.4	5.6
ZnO	235	42.6	21.3	14.2	10.6
Bi₂O₃	235	42.6	21.3	14.2	10.6
BaO	200	50.0	25.0	16.7	12.5
PbO	180	55.6	27.8	18.5	13.9

Table 2: Melting temperature of metals and oxides considered in this work³⁹. Ebullition temperatures are not mentioned because they are beyond the range of interest, dec: decompose.

Compound	Melting temperature (°C)
Rh	1964
Rh ₂ O ₃	1100
RhO	1121
Rh ₂ O	1127
Pt	1768
PtO	507 dec
PtO ₂	450 dec
Pt ₃ O ₄	600 dec
Au	1064
Au ₂ O	200 dec
Au ₂ O ₃	
Al	660
Al ₂ O ₃	2050
Te	449.5
TeO ₂	733
ZnO	1975
ZnF ₂	872
Na ₂ CO ₃	856
Na ₂ O	1134

Table 3: Assignment of absorption bands observed in a 0.2 mm thick 80TeO₂ – 10Na₂O - 10ZnO (at %) glass sample ¹⁹

Band/ μm	Comments	References
≈ 3.03	Stretching mode of free Te–OH groups and/or stretching mode of molecular water	61-63
≈ 3.27	Stretching mode of weak hydrogen-bonded Te–OH groups	61, 63
≈ 4.37	Stretching mode of strong hydrogen-bonded Te–OH groups	61, 63
≈ 6.45	First overtone of O–Te–O symmetric stretch ($2 \times 790 = 1580 \text{ cm}^{-1}$) and/ or bending mode of molecular water	62-64
≈ 7.3	Combination of asymmetric stretch of O–Te–O or Te–O and [TeO ₆] group vibration ($730 + 610 = 1340 \text{ cm}^{-1} = 7.46\mu\text{m}$)	64
≈ 8.26	First overtone of [TeO ₆] group ($2 \times 610 = 1220 \text{ cm}^{-1} = 8.20 \mu\text{m}$)	64
≈ 9.75	Unidentified, possibly first overtone of [TeO ₆] group vibration ($2 \times 490 = 980 \text{ cm}^{-1} = 10.20 \mu\text{m}$)	64

Table 4: Tellurite glass synthesis studies performed under a controlled atmosphere (glove box, quartz reactor...). Synthesis conditions indicate the temperature in °C, the crucible nature (Pt stand for platinum, Au for gold, SiO₂ for quartz, Al₂O₃ for corundum) and the type of mold. Compositions are specified by the metallic element of the corresponding oxide: Te, Zn, W, La, Bi, Nb, Er, Mo, Na, Li, K; OH concentration calculated with $\epsilon(OH) = 47.5 \cdot \ln(10) \cdot L \cdot mol^{-1} \cdot cm^{-1}$ ^{37, 36}.

Purification methods		Results		References
Synthesis temperature / Crucible / mold	Composition	OH Bulk Absorption and residual OH content (ppm wt)	Fiber loss and residual OH content (ppm wt)	
850-900°C & Swirling / Pt or Au & Lid/ Brass	TZNLa	90 dB/m @3.35 μm 6.2 ppm	0.1 dB/m @1.2μm 1.7 dB/m @2.4μm	33
800°C /Pt & Pt stirrer / metallic	TWLa, TP, TG, TZ, TW, TWB, TSbP, TBaSe	-	-	65
Pt-5%Au(wt) or Au & lid / - / -	TZN-Er	-	-	66, 67
800-850°C 3h / Au or Pt / Brass	TZN	-	0.24 dB/m @ 1.55μm	48

Table 5: Tellurite glass purification studies by means of homemade chemicals; all the reported investigations here have been performed under a controlled atmosphere (glove box, quartz reactor...). Purification methods are specified according to the following: PD: Pre-Drying, DR: Desiccant Reagent, HM: Homemade starting products. Synthesis conditions indicate the temperature in °C, the crucible nature (Pt stand for platinum, Au for gold, SiO₂ for quartz, Al₂O₃ for corundum) and the type of mold. Compositions are specified by the metallic element of the corresponding oxide: Te, Zn, W, La, Bi, Nb, Er, Mo, Na, Li, K; OH concentration calculated with $\epsilon(\text{OH}) = 47.5 \cdot \ln(10) \text{ L} \cdot \text{mol}^{-1} \cdot \text{cm}^{-1}$ ^{37, 36}.

Purification methods		Results		References
Purification method / synthesis temperature / Crucible / mold	Composition	OH Bulk Absorption and residual OH content (ppm wt)	Fiber loss and residual OH content	
HM 800°C / Au	TeBiZLi and TeBiZnNa	-	0.05 dB/m @ 1.2 μm	7
PD (500-600° primary vacuum) & HM / 800-850°C, 1-4h, 35g-50g / Pt	TZ and TW, TWLA	0.004 cm ⁻¹ @ 3 μm (0.119 ppm) 0.00117 cm ⁻¹ for driest 0.78T-0.22W (0.035 ppm)	-	16
PD (primary vacuum or O ₂ flux Several hours) & HM & DR for some samples / 800°C / Pt or Au / SiO ₂	TeW, TeWLa, TeWLaBa	0.001 cm ⁻¹ @ 2.95 μm With DR (0.030 ppm)	0.04 dB/m @ 2 μm 0.28 dB/m @ 2.6 μm	17
PD (primary vacuum or O ₂ flux Several hours) & HM & DR / 800°C / Pt or Au / SiO ₂	TeW, TeWLa, TeZnNaBa, TeWLaMo	0.001 cm ⁻¹ @ 2.95 μm With DR (0.030 ppm)	0.2 dB/m @ 2.1 μm	49
HM 800°C stirring / Pt / SiO ₂	TeZn, TeZnNa and TeZnLaNa	0.004 cm ⁻¹ @ 3.4 μm (0.119 ppm)	-	70
HM 800°C stirring / Pt or Au / SiO ₂	TeWLa and TeWBi	TWL 0.15 cm ⁻¹ @ 3 μm (4.48 ppm)		32
HM 800°C stirring / Pt or Au	TeWLa, TeWLaBiEr	0.001 cm ⁻¹ at 3.3 μm (30 ppb)	8dB/m at 3.3 μm Step index fiber (550 ppb wt)	36

DR / 800°C / Pt / SiO ₂	TWLaBi	0.001 cm ⁻¹ @ 2.95 μm (69 ppb)		71
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Table 6: Tellurite glasses purification through the RAP method. All the reported investigations here have been performed under a controlled atmosphere (glove box, quartz reactor...). Purification methods are specified according to the following initials: PD: Pre-Drying, DR: Desiccant Reagent, HM: Homemade starting products, RAP: Reactive Atmosphere Processing. Synthesis conditions indicate the temperature in °C, the crucible nature (Pt stand for platinum, Au for gold, SiO₂ for quartz, Al₂O₃ for corundum) and the type of mold. Compositions are specified by the metallic element of the corresponding oxides: Te, Zn, W, La, Bi, Nb, Er, Mo, Na, Li, K...; OH concentration calculated with $\epsilon(\text{OH}) = 47.5 \cdot \ln(10) \text{ L} \cdot \text{mol}^{-1} \cdot \text{cm}^{-1}$

1 37, 36

Purification methods		Results	References
Purification method / synthesis temperature / Crucible / mold	Composition	OH Bulk Absorption and residual OH content (ppm wt)	
<p>PD (vacuum 100-300°C) under the RAP (O₂-CCl₄) or DR (PbCl₂ substituted by PbF₂) or Addition of NH₄NO₃, CaCl₂, AlF₃ / 800°C 30' O₂/ Pt, Au, SiO₂ or Al₂O₃</p>	TeZn and TePbPbCl ₂	At 3 μm, 0.36 cm ⁻¹ to 0.16 cm ⁻¹ when Cl ₂ atmosphere is used. (10.74 to 4.77 ppm)	51
<p>PD (623-723K -under O₂-Cl₂ 3h 0.5L.min⁻¹) RAP (Cl₂-O₂ (10-90% vol) / 1073K O₂-Cl₂ / Au / Brass</p>	80Te10Zn10Na	0.06 cm ⁻¹ @ 3μm (1.79 ppm)	46
<p>RAP (CCl₄/O₂) / 800-1100°C / Au & Al₂O₃</p>	TeGeZnYNaEr	1.5 cm ⁻¹ @ 3μm 44.8 ppm	47
<p>PD (15h T<Tf starting product + all batch under vacuum at 140°C) RAP Ar/F₂ 6h @600°C / 1400-1575°C-O₂-1h / Pt / C</p>	P ₂ O ₅ -B ₂ O ₃ -Nb ₂ O ₅	1 cm ⁻¹ @3μm (29.8 ppm)	29

Table 7: Tellurite glass purification by desiccant reagents studies. All the reported investigations here have been performed under a controlled atmosphere (glove box, quartz reactor...). Purification methods are specified according to the following initials: PD: Pre-Drying, DR: Desiccant Reagent, HM: Homemade starting products. Synthesis conditions indicate the temperature in °C, the crucible nature (Pt stand for platinum, Au for gold, SiO₂ for quartz, Al₂O₃ for corundum) and the type of mold. Compositions are specified by the metallic element of the corresponding oxides: Te, Zn, W, La, Bi, Nb, Er, Mo, Na, Li, K...; OH concentration calculated with $\epsilon(OH) = 47.5 \cdot \ln(10) \text{ L} \cdot \text{mol}^{-1} \cdot \text{cm}^{-1}$ ^{37, 36}.

Purification methods		Result		References
Purification method/synthesis temperature/Crucible/mold	Composition	OH Bulk Absorption and residual OH content (ppm wt)	Fiber loss and residual OH content (ppm wt) ^b	
DR (ZnF ₂ , KCl, NH ₄ F-HF, ZnO replaced by ZnF ₂ (20% at) with or a without NH ₄ F-HF heat treatment K or Li/15h 500°C for decarbonation / 45'500°C ; 850°C 5-10' O ₂ / Pt / -	TeBiZn + (Ba/K/Li)	<0.5 cm ⁻¹ (14.6ppm with ZnF ₂ and NH ₄ F-HF - < 14.9 ppm) ZnF ₂ only: 1 cm ⁻¹ (29.8 ppm)	Background improvement of O ₂ rich atmosphere between 0.8 and 1.6 μm	20
PD (525-750°c vacuum) & DR (TeCl ₄) / 800°C (optimized Thermal profile) / Pt / Pt	TeW	7.5 dB/m (0.017 cm ⁻¹) @ 3 μm (0.51 ppm)		74
DR (NaCl and PbCl ₂ up to 20% at) / 70g,700-900°C / Au / metal	TeZnNa & TePbPbCl ₂	0.3 cm ⁻¹ @ 3.35μm (8.9ppm) 0.16 cm ⁻¹ @ 3.35μm (4.8 ppm)	45dB/m (3 ppm) 13.5dB/m (0.93 ppm)	60
DR / 800°C / Pt / SiO ₂	TWLaBi	0.001 cm ⁻¹ @ 2.95 μm (0.069 ppm)		71
DR (ZnF ₂) / 750°C 45' & 800°C 15' 650°C for casting / Pt + lid / Aluminum	TeZnNa	20 dB/m @3.35 μm (1.37 ppm)		19
DR (ZnF ₂) / 730-825°C 30' & Stirring / Pt lid	TZnF ₂ PbNb 0% ZnO (only use of ZnF ₂)	0.15 cm ⁻¹ @3 μm (4.48 ppm)		75
DR (ZnF ₂ up to 5 % at) / 750-950°C Stirred & bubbled with flowing gas / Au / -	TeZnLa	0.003 cm ⁻¹ @ 3. μm. (0.21 ppm)	<0.5 dB/m up to 4.2μm 0.25 dB/m @ 3μm (0.017ppm)	11, 30
DR (ZnF ₂ up to 30% at) / 800°C-2h	TeZnNa	0.03cm ⁻¹ @ 3 μm (2 ppm)		52
DR (ZnF ₂) / 750-850°C 30', 700°C casting / Au / Graphite	TeWZn			53
PD (primary vacuum)	TeW, TeWLa,	0.001 cm ⁻¹ @ 3μm	<0.6dB/km between 1.4 et 2.4	49

^b According to ^{60, 60}, OH absorption difference between bulk and fiber may be partly due to residual water adsorbed on bulk surfaces. Churbanov et al ^{16, 16} and Anashkina et al ^{36, 36} show that it is necessary to account for this effect for short samples presenting low OH quantity.

Or O ₂ flux several hours) & DR (BiF ₃ , WOF ₄) / 800°C / Pt or Au / -	TeZnNaBi, TeWLaMo	(0.030 ppm) 74 dB/km @ 1.55μm	μm	
DR (BaCl ₂)	TeBaZn	8 dB/m @ 3.35 μm (0.55 ppm)		76
PD (35 min - 130°C under vacuum) & DR (BaF ₂ 19-38% at) / 900°C 35' / Al ₂ O ₃ / Copper	TeBaF ₂ NaEr (0% BaO) only BaCl ₂	0.0702 cm ⁻¹ @ 2.7μm		77
PD (35 min - 130°C under vacuum) & DR (BaF ₂ 25% at) / 900°C 35' / Al ₂ O ₃ / Copper	TeBaF ₂ YHo & TeAlBCaYSrMg, all fluoride except for TeO ₂ (fluorotellurite glass)	Flat transmission < 1dB/m @ 4μm 0.0006cm ⁻¹ between 3 et 4μm (0.018 ppm)		78, 79
PD (vacuum) & DR (ZnF ₂ , PbCl ₂ , BaF ₂ , PbF ₂ and CaF ₂) / 700-800°C / O ₂ bubbling / Au / Stainless steel	TeZnNaLa, TeZnLa with PbCl ₂ & NaF & PbF ₂ & CaF ₂ TeZnNb, TeZnW, TeZnMo with ZnF ₂	0.054 cm ⁻¹ @ 3 μm (1.6 ppm)		50
DR (ZnF ₂ >10% at) Room atmosphere / 750-850°C / Au / Graphite	TeWZnF ₂			57
DR (ZnCl ₂ >20% at) Room atmosphere / 800-1000°C / Al ₂ O ₃ / Brass	TeZnZnCl ₂			59
DR (ZnF ₂ 5-10% at) / 850°C O ₂ / Pt / Brass	TeZnNa	0.008 cm ⁻¹ @ 3.4 μm (0.24 ppm)	5% at ZnF ₂ 14dB/m @ 3.4 μm (0.96 ppm)	21
DR (ZnF ₂ 5-10% at) / 850°C O ₂ / Pt / Brass	TeZnNa		5% at ZnF ₂ Transmission on fiber up to 4 μm 5dB/m @ 3.5 μm (0.34 ppm)	80
DR (ZnF ₂ 5-10% at) / 850°C O ₂ / Pt / Brass	TeZnNa/TeNaGeZ		3dB/m @ 3.5 μm (0.206 ppm)	81, 27
DR (BaF ₂ 20% at) & fluoride cladding / - N ₂ / - / -	TeBaF ₂ Y ₂ O ₃ /TeAl BaCaYSrMg all fluoride except for TeO ₂ (fluorotellurite glass)		Core 20% at BaF ₂ Clad 90% at fluoride <2.3 dB/m from 1.1 to 4.5 μm (0.07ppm)	82

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Figure 1: (a) Ellingham diagram of Pt, Au, Te and Al for the reactions of oxidation;

(b) $\log(P_{O_2}/P^\circ)$ at the equilibrium versus the temperature.

Figure 2: Transmission spectra of two tellurite glasses of the same composition, prepared in gold or platinum crucible.

Figure 3: Comparison between the absorption coefficient of a tellurite glass made under ambient atmosphere (TZN air), the same glass prepared under a pure O_2 atmosphere (TZN O_2)²¹ and the same composition made under a pure O_2 associated to optimized time and temperature conditions.³³ The TZN glass composition is $80TeO_2-10ZnO-10Na_2O$ (mol %).

Figure 4: (a) OH absorption coefficient for different bulk tellurite glasses made from homemade ultra-pure starting products and (b) fibers attenuation.

Figure 5: Ellingham diagrams of the reaction between Cl_2 and hydrated starting products of TZN glasses; considered reactions are specified in the legend and are written for one mole of Cl_2 .

Figure 6: Bulk coefficient absorption function of wavelength for tellurite glasses prepared through the RAP

Figure 7: (a) The Ellingham like diagram of the reaction between ZnF_2 and hydrated starting products, considered reactions are specified in the legend. (b) comparison of the Gibbs free energy of the reactions of water with TeO_2 and ZnF_2 . (c) and (d) are the log of the equilibrium constant's plot function of the temperature corresponding to the cases described in figure (a). The region where products or reagents are favored is specified on the figures.

Figure 8: Bulk absorption coefficient as a function of wavelength for tellurite glasses prepared through reagent addition: (a) results showing a low efficiency of the method; (b) moderate efficiency; (c) high efficiency and best results reported, OH quantity is strongly reduced, this absorption level allows for transmission on meter long samples i.e. fibers.

Figure 9: Reported attenuation measurements recorded by means of the cut back method on fiber samples (length >50cm) made from glasses prepared through reagent addition, (a) results obtained by different research groups, (b) results we achieved at ICB.

Figure 10: Broadest SC spectra generated in tellurite fibers: green curve is the SC presented in ^{11, 22} has been obtained by coupling 94 nJ pulse energy centered at 1.92 μm into a fiber showing a W refractive index profile, using a 26 ps source operating at 32 MHz ; in red the SC obtained by a collaboration between University of Bourgogne and University of Stuttgart ²⁷ has been generated in the step index tellurite fiber which attenuation is shown in figure 9 (orange curve) ²⁸ by injecting 3.5 nJ pulse energy at 2.4 μm using a 350 fs source operating at 43 MHz and in black the SC generated by pumping at 3 μm with 1.4 nJ coupled energy in the same fiber than the one which leads to the green SC but with the same pump source as the one used in the case of the red SC i.e. 350 fs, 43MHz ²⁶. The blue SC is generated by

launching a 196 fs pump - 31 nJ pulse centered at 2 μ m in a fluorotellurite fiber⁸⁵ whose attenuation is shown by the blue curve in figure 9(a). For each of the spectra, the 0-dB level corresponds to the maximum detected intensity which typically lies at the pump wavelength.