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1 **Safer and stronger together? Effects of the agglomeration on nanopowders**  
2 **explosion**

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11

12 **Abstract:**

13 Among the factors influencing dust explosion, the particle size distribution (PSD) is both one  
14 of the most important and complex to consider. For instance, it is commonly accepted that the  
15 explosion sensitivity increases when the particle size decreases. Such an assertion may be  
16 questionable for nano-objects which easily agglomerate. However, agglomerates can be  
17 broken during the dispersion process. Correlating the explosion parameters to the actual PSD  
18 of a dust cloud at the moment of the ignition becomes then essential. The effects of the  
19 moisture content and sieving were investigated on a nanocellulose powder and the impact of a  
20 mechanical agglomeration was evaluated using a silicon coated by carbon powder. Each  
21 sample was characterized before and after dispersion using in situ laser particle size  
22 measurement and a fast mobility particle sizer, and explosion and minimum ignition energy

23 tests were conducted respectively in a 20 L sphere and in a modified Hartmann tube. It was  
24 observed that drying and/or sieving the nanocellulose mainly led to variations in terms of  
25 ignition sensitivity but only slightly modified the explosion severity. In contrast, the  
26 mechanical agglomeration of the silicon coated by carbon led to a great decrease in terms of  
27 ignition sensitivity, with a minimum ignition energy varying from 5 mJ for the raw powder to  
28 more than 1J for the agglomerated samples. The maximum rate of pressure rise also decreased  
29 due to modifications in the reaction kinetics, inducing a transition from St2 class to St1 class  
30 when agglomerating the dust.

31 **Keywords:** Dust explosion; Nanoparticles; Agglomeration; Dispersion

32

33

## 34 **1. Introduction**

35 According to the European Commission (2011/696/EU, 2011), a nanomaterial is “a natural,  
36 incidental or manufactured material containing particles, in an unbound state or as an  
37 aggregate or as an agglomerate and where, for 50% or more of the particles in the number size  
38 distribution, one or more external dimensions is in the size range 1 nm – 100 nm”. This  
39 recommendation also specifies that any material presenting a specific surface area by volume  
40 of the material higher than  $60 \text{ m}^2 \cdot \text{cm}^{-3}$  must be considered as a nanomaterial. This small size  
41 induces additional, enhanced or different properties for the nanomaterials, which implies that  
42 the fundamental properties such as chemical, mechanical, optical or biological properties are  
43 modified and often innovative (Stark et al., 2015). Those new properties resulted to an  
44 increasing interest in nanoparticles for applications in different fields, like health, automotive  
45 industry, construction, food or electronic sector, but also generated a question concerning the

46 toxicity of the materials (Oberdörster et al., 2005). But nanoparticles, as well as any  
47 combustible particles, present another important risk: their dispersion under certain conditions  
48 and in the presence of an ignition source can lead to an explosion.

49 If dust explosion risks concerning micron-sized particles can be relatively well estimated, it is  
50 not yet the case for nanoparticles. Indeed, the interest in nanotechnologies and in their specific  
51 properties is quite recent and there are only a few feedbacks on accidental dust explosions  
52 involving nanoparticles (Wu et al., 2014). Some laboratory tests were already carried out to  
53 evaluate the ignition sensitivity and explosivity of nanopowders (Boillard et al., 2013;  
54 Holbrow et al., 2010; Krietsch et al., 2015; Wu et al., 2009). It appears that powders are  
55 usually more sensitive to ignition, but no significant variation concerning the explosion  
56 severity was observed (Bouillard et al., 2010; Dufaud et al., 2011; Holbrow et al., 2010). This  
57 phenomenon is mainly explained by the agglomeration and aggregation phenomenon induced  
58 by the small size of the nanoparticles, which reduces the reactive surface area and decreases  
59 the explosion severity (Eckhoff, 2011, 2012).

60 Generally, a nanopowder is comprised of primary nanoparticles, i.e. individual nanoparticles,  
61 which aggregate, forming so-called “primary aggregates”, which themselves agglomerate  
62 with each other. An agglomerate consists of weakly bonded particles that can be separated  
63 while an aggregate is an assembly of strongly bonded particles that cannot be broken  
64 (Sokolov et al., 2015; Walter, 2013). The main cohesion forces inducing the agglomeration of  
65 nanoparticles aggregates are the van der Waals interactions, electrostatic and magnetic forces,  
66 and, in the case of wet particles, capillary effects (Debrincat et al., 2008; Hartley et al., 1985).  
67 Each agglomerate then possesses a cohesion strength that can be calculated by different ways  
68 (Kendall, 1988; Rumpf, 1962; Weiler et al., 2010). In order to break an agglomerate, an  
69 energy higher than the cohesion strength must be provided, leading back to the primary  
70 aggregates.

71 For an explosion to occur, some conditions are needed. Among them is the dispersion of the  
72 powder, which can break the agglomerates, at least partially, depending on the shear stress  
73 occurring during the dust dispersion. Since the particle size distribution (PSD) is a very  
74 important parameter influencing dust explosion, it becomes necessary to characterize the dust  
75 cloud not only before its dispersion, but also at the exact ignition time (Santandrea et al.,  
76 2019b). Moreover, it is imperative to choose wisely the operating conditions for the explosion  
77 tests, to be sure to evaluate the worst -but realistic- case scenario. Indeed, the standard  
78 conditions were established for micron-sized particles and are currently applied when testing  
79 nanoparticles. However, these conditions may need to be adapted due to the specific  
80 properties of nanomaterials (Santandrea et al., 2020).

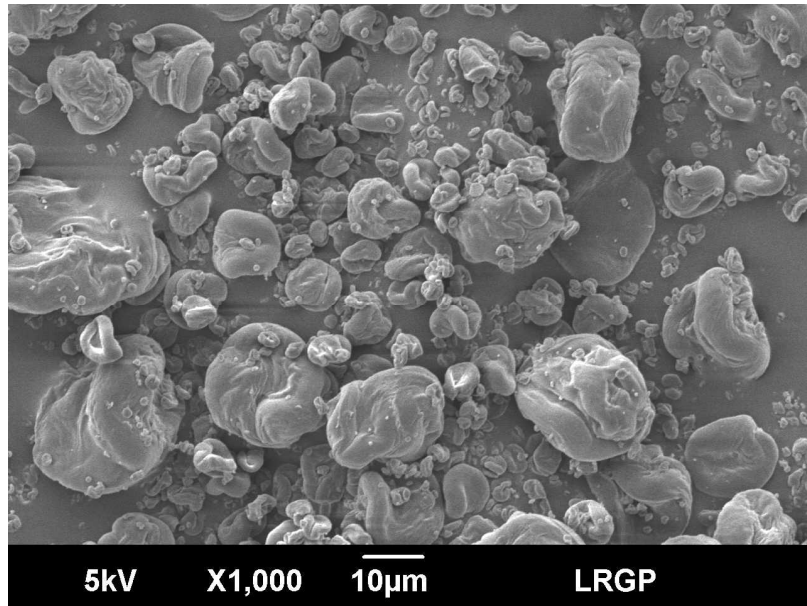
81 The impact of the agglomeration on the explosion severity of nanopowders was evaluated by  
82 modifying the agglomerates size and cohesion strength before their dispersion and performing  
83 explosion tests in a 20 L sphere according to international standards (EN 14034-1, 2004; EN  
84 14034-2, 2006). Since the addition of a cohesive agent would modify the reactions involved  
85 in the explosion, three main ways can be considered: grinding, selection (sieving), and  
86 mechanical agglomeration (e.g. wet or dry granulation, compaction). The dispersion  
87 procedure in the 20 L sphere already inducing a high shear stress due to the pressurization of  
88 the dust container at 20 barg, it would be difficult to further reduce the size of the  
89 agglomerates after dispersion. Indeed, the smaller the particle (or agglomerate), the harder it  
90 is to be broken (Deng et al., 2016). Moreover, dry powder grinding would not allow to obtain  
91 agglomerates smaller than around 1  $\mu\text{m}$ , and the process would produce heat and possibly  
92 electrostatic discharges that could be sufficient to ignite the most sensitive powders.  
93 Therefore, powder grinding was not considered as a suitable solution. The effects of the  
94 agglomeration on the explosion severity and ignition sensitivity were then investigated by

95 sieving and by mechanical agglomeration, respectively on nanocellulose and carbon coated  
96 silicon powders.

## 97 **2. Materials and methods**

### 98 **2.1. Materials**

99 The effects of nanopowders agglomeration on their explosivity were investigated using two  
100 different materials. First, a nanocrystalline cellulose (NCC from CelluForce), called  
101 ‘nanocellulose’, was chosen due to its organic nature and wide range of applications. This  
102 powder is constituted of primary nanofibers of 3 nm width and an average length of 70 nm  
103 which form agglomerates with diameters ranging between 1  $\mu\text{m}$  and several dozens of  
104 micrometers, as presented in Figure 1. To avoid the influence of humidity on both  
105 agglomeration state (capillary effects) and explosion characteristics, which was discussed by  
106 Santandrea et al. (2020), this powder was systematically dried at 90°C under vacuum. The  
107 water activity, i.e. the partial vapor pressure of water divided by the standard state partial  
108 vapor pressure of water, was measured around 0.03 after drying (Aw-meter, Rotronic),  
109 confirming that no additional water present in the powder would participate to the  
110 combustion. In an attempt to obtain different agglomerate sizes, the nanocellulose was sieved  
111 using a 40  $\mu\text{m}$  mesh strainer submitted to low amplitude vibrations to avoid the breakage of  
112 the agglomerates. This procedure then allows the separation of the biggest agglomerates from  
113 the finest particles.



114

115

*Figure 1. Raw nanocellulose observed by Scanning Electron Microscopy*

116

Complementary to the PSD selection performed on nanocellulose, the effects of a mechanical

117

agglomeration without any additive were investigated on carbon coated silicon, noted 'Si $\Omega$ C',

118

through different agglomerated samples directly supplied by Nanomakers. Due to the low

119

density of the nanoparticles, such modification notably facilitates the transportation of the

120

powder and increases the powder flowability. Contrary to the nanocellulose, the powder was

121

not dried to avoid the breakage of the agglomerates, as the mechanical agglomeration was

122

realized directly on the raw powder without any additive. Moreover, the oxidation of silicon

123

in the presence of water can lead to the formation of hydrogen (Mehta et al., 2014), which

124

would increase the explosion severity. In this case, drying the powder would probably lead to

125

an underestimation of the safety parameters of the powders.

126

Four samples of carbon coated silicon constituted of the same powder with different densities,

127

agglomerates sizes and cohesion strengths were then studied in this work: the raw powder of

128

primary diameter of 40 nm and density of 40 g.L<sup>-1</sup>, two samples agglomerated according to a

129

process noted 'process A' with bulk densities of respectively 260 g.L<sup>-1</sup> and 400 g.L<sup>-1</sup>, which

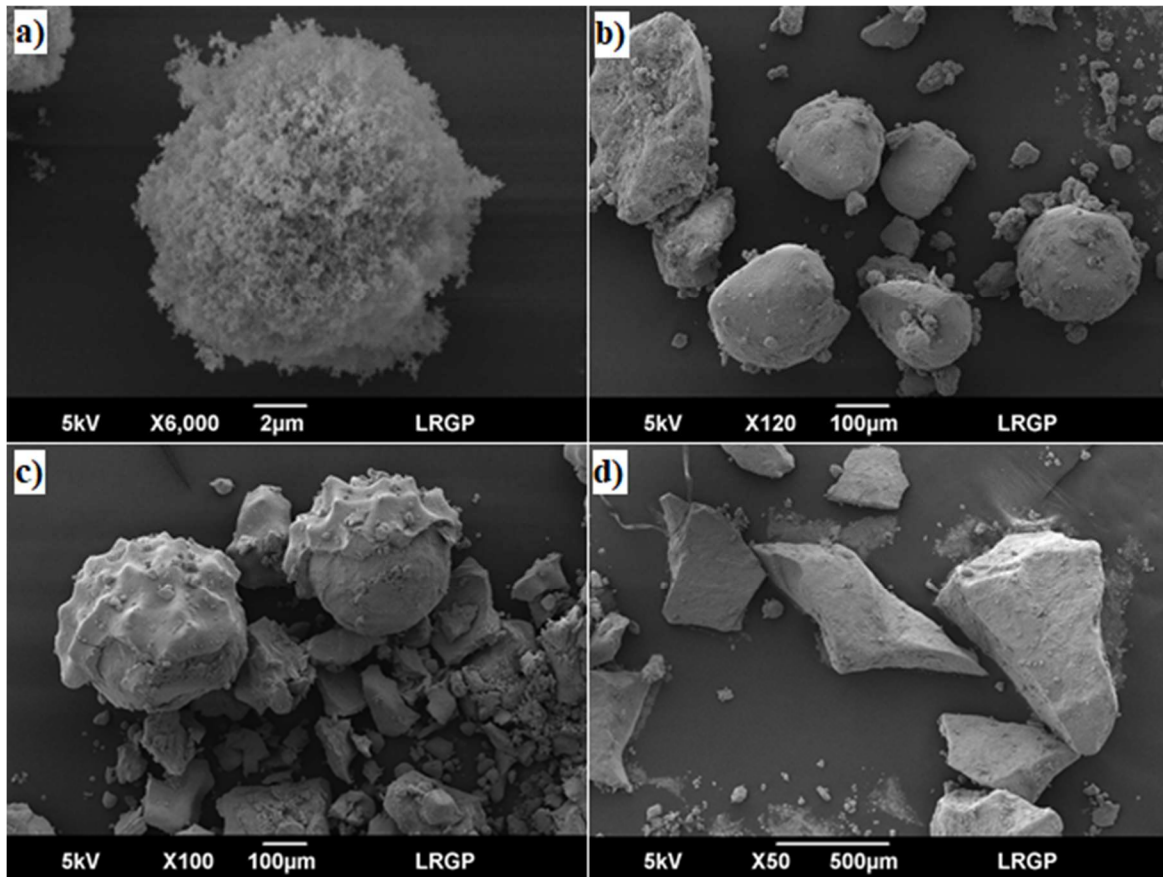
130

will be noted powders A1 and A2, and a sample agglomerated according to a 'process B' with

131 a bulk density of  $400 \text{ g.L}^{-1}$ , noted sample B. The specific processes of agglomeration are  
132 confidential and unfortunately cannot be described here. Scanning Electron Microscopy  
133 analyses were performed on each sample, without any modification, to visualize the shape  
134 modifications due to the agglomeration (Figure 2).

135 The raw powder appears to be constituted of small spherical agglomerates of apparent  
136 diameters lower than  $10 \mu\text{m}$ , as shown in Figure 2a. The modified powders also present the  
137 same kind of agglomerated structures, but also bigger agglomerates. Indeed, the powder A1  
138 seems essentially composed of big 'roughly spherical' agglomerates with a diameter around  
139  $100$  to  $200 \mu\text{m}$  (Figure 2b). Similar agglomerates can be found in the sample A2, but with  
140 agglomerates size around  $300 \mu\text{m}$  (Figure 2c). Finally, when using the process B to  
141 agglomerate the powder, non-spherical agglomerates with an average size reaching  $500 \mu\text{m}$   
142 were formed (Figure 2d).





143

144 *Figure 2. Carbon-coated silicon observed by Scanning Electron Microscopy a) Raw powder*

145 *b) Sample A1 (260 g.L<sup>-1</sup>) c) Sample A2 (400 g.L<sup>-1</sup>) d) Sample B (400 g.L<sup>-1</sup>)*

146 **2.2. Methods**

147 The initial particle size distribution (PSD), i.e. before dispersion in the testing equipment, of  
 148 each sample of nanocellulose and carbon coated silicon was measured in air using a laser  
 149 diffraction HELOS/KR-Vario (Sympatec GmbH). The PSD was characterized by  
 150 sedimentation of the powders in the measuring area of the apparatus. As this process was  
 151 realized manually (by sprinkling powders with a spatula shaken at a constant height), it may  
 152 therefore be questioned in terms of repeatability. Nevertheless, several tests on each sample  
 153 led to similar PSD, which then gives a good order of magnitude of the PSD of the dust cloud  
 154 submitted to very low shear rates. In this study, the surface fraction was considered instead of  
 155 the ‘commonly-used’ volume fraction when analyzing the PSD in order to highlight the

156 surface specificities of nanoparticles. Moreover, when considering the volume fraction, the  
157 high volume of big agglomerates tends to occult the presence of smaller nano-agglomerates.  
158 Thus, the volume fraction measured by the apparatus was converted into a surface fraction  
159 using the equivalent volume diameter.

160 Since the dispersion procedure, required to produce a dust explosion induces a shear stress  
161 that can potentially break the agglomerates, PSD measurements were conducted after  
162 dispersion in the explosion vessel. However, no apparatus or technique currently allow the  
163 determination of a PSD over a wide range from 10 nm (primary particles) to 500  $\mu\text{m}$   
164 (agglomerates), at high concentration (above the lower explosion limit or at least at a few  $\text{g}\cdot\text{m}^{-3}$ )  
165 and at high frequency (each millisecond or at least each 10 ms). Thus, different techniques  
166 were combined. The dust dispersion was realized in a 20 L sphere equipped with windows  
167 according to the same procedure than during explosion tests: the weighed sample is placed in  
168 the dust container, and the sphere is evacuated to 0.4 bara. The container is then pressurized  
169 20 barg and the electrovalve connecting the dust canister to the explosion chamber opens,  
170 inducing the dispersion of the powder. Contrary to explosion tests, no ignition was performed,  
171 and the time evolution of the PSD was recorded.

172 The laser diffraction sensor (Helos - Sympatec) used to measure the initial PSD of the dust  
173 before dispersion was attached to the 20 L sphere equipped with visualization windows made  
174 of borosilicate with a diameter of 9.7 cm to allow optical measurements (Murillo, 2016). The  
175 PSD was then measured at the place of ignition using various lenses (called R1, R3 and R5)  
176 and systematically presented at the moment of ignition, i.e. 60 ms after the beginning of the  
177 dispersion according to EN 14034 1&2 (2004; 2006). The boundaries of the measurement  
178 ranges are 0.1 – 35  $\mu\text{m}$ , 0.5 – 175  $\mu\text{m}$  and 0.5 – 875  $\mu\text{m}$  for R1, R3 and R5 lenses,  
179 respectively. It should be noted that the maximum distance between the lens and the sample is  
180 20 mm for R1, whereas it can reach 47 cm for R5. As a consequence, only the powder located

181 near the observation windows will be analyzed by using R1, whereas the other lenses will  
182 give a PSD representative of the overall content of the sphere.

183 The presence of nanoparticles after dispersion of nanocellulose and carbon coated silicon in  
184 the 20 L sphere was also investigated using a Fast Mobility Particle Sizer (FMPS - TSI)  
185 measuring electrical mobility diameters from 5.6 to 560 nm with 1 Hz time resolution and a  
186 Scanning Mobility Particle Sizer (SMPS), providing one measurement every two minutes.  
187 Despite the low frequency and mass concentration which prevent an accurate characterization,  
188 the observation of nanoparticles up to two minutes after dispersion, with particle size  
189 distribution ranging from 10 to 400 nm, implies that such particles were also present a few  
190 dozens of milliseconds after dispersion.

191 The explosion severity of each sample was measured in a standard 20 L sphere according to  
192 EN 14034 1&2 (2004; 2006), i.e. using two chemical igniters of 5 kJ each and an ignition  
193 delay time,  $t_v$  of 60 ms. The 20 L sphere is equipped with a cooling jacket with water at 300  
194 K. The minimum ignition energy (MIE) of the samples was measured using a standard  
195 modified Hartmann tube according to ISO/IEC 80079-20-2 (2016). Although international  
196 standards EN 14034-3 (2006) recommend to measure the lower explosion limit in the 20 L  
197 sphere using an ignition energy of 2 kJ, this parameter was approximated in the same  
198 conditions as the explosion severity tests, i.e. for an ignition energy of 10 kJ.

### 199 **3. Results and discussions**

#### 200 **3.1. Effects of particle size selection: example of nanocellulose**

201 As detailed in section 2.1, the nanocellulose was dried and sieved using a 40  $\mu\text{m}$  mesh  
202 strainer. The particle size distribution of each fraction was measured using the laser  
203 diffraction sensor (R3 lens) by sedimentation. The upper fraction was constituted of  
204 agglomerates of sizes between 40 and 100  $\mu\text{m}$ , whereas the lower fraction exhibited a surface

205 diameter around 10  $\mu\text{m}$ . The agglomerates sizes of each sample are summarized in Table 1. It  
 206 should be noted that the sieving process does not modify the size of the agglomerates and  
 207 only aims at separating the agglomerates according to their sizes. However, the biggest  
 208 agglomerates, i.e. the less cohesive, may be broken during the sieving process (Deng et al.,  
 209 2016).

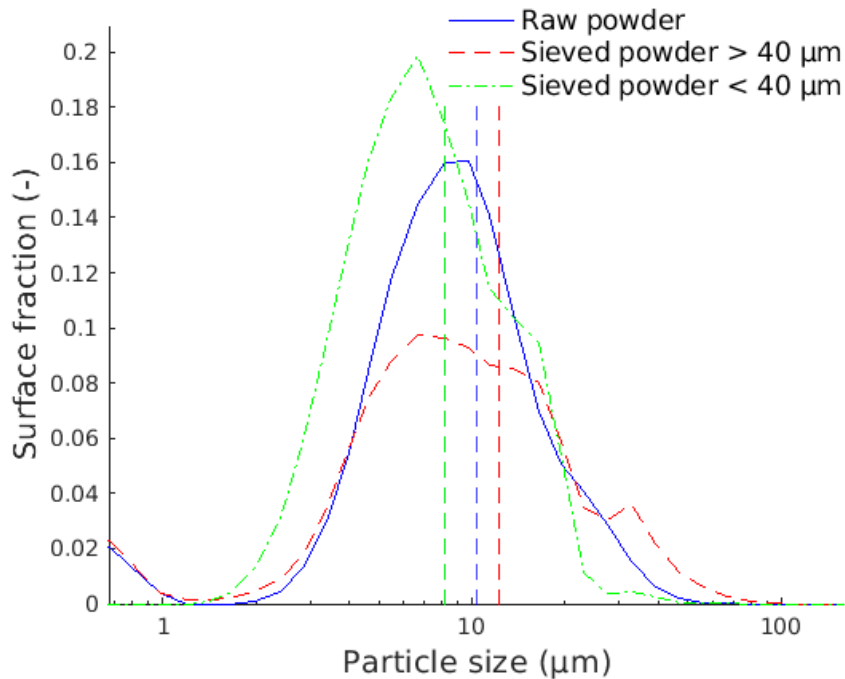
210 *Table 1. Characteristics of the raw and sieved nanocellulose*

Powder	Raw powder	Fine fraction	Large fraction
Sieving size ( $\mu\text{m}$ )	N/A	< 40	> 40
Mean surface diameter by sedimentation ( $\mu\text{m}$ )	48	10	54
Mean surface diameter 60 ms after dispersion in the 20 L sphere ( $\mu\text{m}$ )	10	8	12

211

212 The dust clouds produced after dispersion in the 20 L sphere were also characterized at the  
 213 moment of ignition, i.e. 60 ms after the dust injection (Table 1). First, it can be confirmed that  
 214 the injection system tends to break the powder agglomerates, as already shown by previous  
 215 studies (Du et al., 2015; Sanchirico et al., 2015). Despite the high shear stress induced by the  
 216 dispersion system of the 20 L sphere, a slight difference in the particle size distributions of the  
 217 samples can be observed in Figure 3. Indeed, the mean surface diameter obtained by  
 218 dispersion of the powder sieved below 40  $\mu\text{m}$  is 8  $\mu\text{m}$  whereas the dispersion of the bigger  
 219 agglomerates (> 40  $\mu\text{m}$ ) led to a mean surface diameter of 12  $\mu\text{m}$ . As a comparison, the mean  
 220 surface diameter by dispersion of the raw powder is 10  $\mu\text{m}$ , which confirms the efficiency of  
 221 the sieving procedure.

222



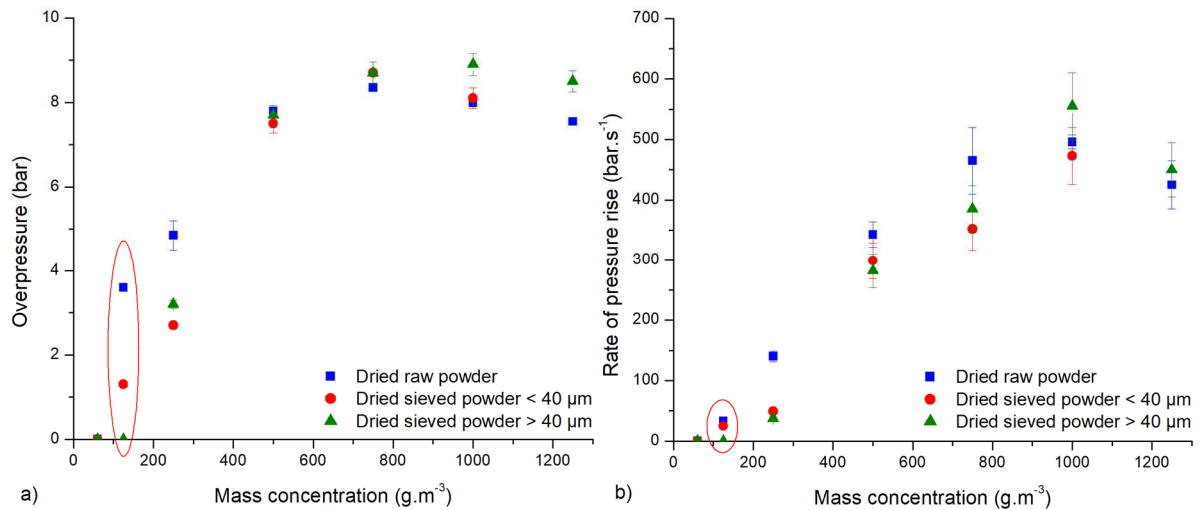
223

224 *Figure 3. Particle size distribution of raw and sieved nanocellulose after injection in the 20 L*  
 225 *at the place and moment of ignition (60 ms after dispersion) measured by the Helos*  
 226 *diffraction sensor (R3 lens, 0.5 – 175 μm)*

227

228 The explosion tests performed on the different samples conducted to rather similar explosion  
 229 severities, probably due to the small differences in the PSD after dispersion. Indeed, a  
 230 maximum overpressure of  $8.9 \pm 0.4$  bar and a maximum rate of pressure rise of  $555 \pm 66$   
 231  $\text{bar}\cdot\text{s}^{-1}$  were obtained for the bigger agglomerates ( $> 40 \mu\text{m}$ ) whereas these values respectively  
 232 reached  $8.7 \pm 0.4$  bar and  $473 \pm 56 \text{ bar}\cdot\text{s}^{-1}$  for the finest powder (Figure 4), which does not  
 233 constitute a significant difference with regard to the experimental uncertainties. Nevertheless,  
 234 the lowest concentration inducing an explosion was obtained at  $125 \text{ g}\cdot\text{m}^{-3}$  with the raw  
 235 powder and the smallest agglomerates and only at  $250 \text{ g}\cdot\text{m}^{-3}$  with the sample of agglomerates  
 236 bigger than  $40 \mu\text{m}$ . Thus, despite the breakage of the majority of the agglomerates observed in  
 237 Figure 3, the sample constituted of agglomerates bigger than  $40 \mu\text{m}$ , which contains fewer  
 238 fine particles than the initial powder, appears to be less ignition sensitive than the two other

239 samples. This observation tends to corroborate the theory stipulating that the ignition  
240 sensitivity of a powder is mainly affected by the presence of fine particles (Saeed et al., 2019)  
241 as they are more prone to volatilize.



243 *Figure 4. Effects of nanocellulose sieving on a) the maximum overpressure and b) the*  
244 *maximum rate of pressure rise*

245  
246 To further investigate on the influence of fine particles on the ignition sensitivity, ignition  
247 energy tests were conducted in the modified Hartmann tube. The minimum ignition energy of  
248 the raw dried powder was evaluated at 5 mJ. After drying and sieving at 40  $\mu\text{m}$ , both obtained  
249 fractions also presented minimum ignition energies of 5 mJ. As their MIE is lower than  
250 10 mJ, these powders can be considered as being very sensitive to electrostatic ignition (Janès  
251 et al., 2008). However, when drying and sieving at 70  $\mu\text{m}$ , the bigger fraction exhibited a  
252 minimum ignition energy of 14 mJ. It should be noted that the MIE of microcrystalline  
253 cellulose (MCC) ranges from 30 to 100 mJ for 30  $\mu\text{m}$  particles (BGIA, 1997) and previous  
254 tests performed on MCC with a mean diameter of 100  $\mu\text{m}$  led to a MIE of 590 mJ .  
255 These tests highlight the effects of the presence of fine particles on the ignition sensitivity,  
256 especially for nanomaterials comprised of various sizes of agglomerates. Nevertheless,

257 sieving is probably not the most accurate method of powder selection, as small particles can  
258 stick to agglomerates (due to electrostatic interactions, for instance) and remain in the upper  
259 fraction, although smaller than the mesh. Furthermore, for non-spherical particles, only one  
260 dimension smaller than the mesh is sufficient for an agglomerate to be retrieved in the low  
261 fraction, independently from its volume. As a consequence, in order to stress the influence of  
262 agglomeration of nanopowders on their explosion severity and MIE, the impact of mechanical  
263 agglomeration has been studied.

264

### 265 **3.2. Effects of mechanical agglomeration: example of carbon coated silicon**

266 Since the agglomerates that naturally exist in the raw powder (especially for nanocellulose)  
267 are easily breakable and in order to assess the influence of the particle size distribution on the  
268 explosion severity, the effects of mechanical agglomeration were investigated through carbon  
269 coated silicon (Si $\Omega$ C). The initial mean surface diameter was measured by sedimentation of  
270 the powder using the laser diffraction sensor (Helos – Sympatec) and by wet dispersion in  
271 ethanol using a Mastersizer 2000 S (Malvern Instruments). Analyzing the wet dispersion  
272 measurements presented in Table 2, it appears that all the agglomerated samples present a  
273 mean diameter much higher than that of the raw powder, i.e. 68  $\mu\text{m}$ , and that the powder B  
274 exhibits the highest diameter, reaching 442  $\mu\text{m}$ , which is consistent with the SEM  
275 observations. However, when regarding the powders agglomerated by process A, it seems that  
276 the powder A1 (260  $\text{g.L}^{-1}$ ) is more agglomerated than the powder A2 (400  $\text{g.L}^{-1}$ ), which seems  
277 inconsistent with SEM observations. Once again, it should be kept in mind that the wet  
278 dispersion in ethanol modifies the interactions between particles (variation of zeta potential)  
279 and does not provide accurate information concerning the particle size distribution of a  
280 powder in air.

281 The surface diameter was measured by sedimentation of the powder using two lenses of the  
 282 laser diffraction sensor: R3 and R5 (Table 2). Tests performed with the R3 lens led to rather  
 283 similar mean surface diameters by sedimentation, from 20  $\mu\text{m}$  to 32  $\mu\text{m}$ , which can be due to  
 284 the upper limitation of the lens. The R5 lens allows the measurement of particles from 4.5 to  
 285 875  $\mu\text{m}$ , inducing that the agglomerates observed by SEM can be distinctly observed. First,  
 286 the powder B is then confirmed to be the most agglomerated ( $d_{50,R5} = 297 \mu\text{m}$ ), although this  
 287 was not visible using the R3 lens. Then, the sample A2 presents a lower mean diameter than  
 288 the sample A1, which is conversely proportional to the particle density and confirms the  
 289 measurements performed by wet dispersion. Finally, the raw powder presents a mean surface  
 290 diameter of 207  $\mu\text{m}$ , which is considerably higher than the size of the agglomerates observed  
 291 by SEM. It should be reminded that this lens does not allow the measurement of particles  
 292 smaller than 4.5  $\mu\text{m}$ , implying that the small agglomerates may not be measured by the  
 293 apparatus.

294

295 *Table 2. Mean diameter of the different samples of carbon-coated silicon measured by*  
 296 *sedimentation and wet and dry dispersion*

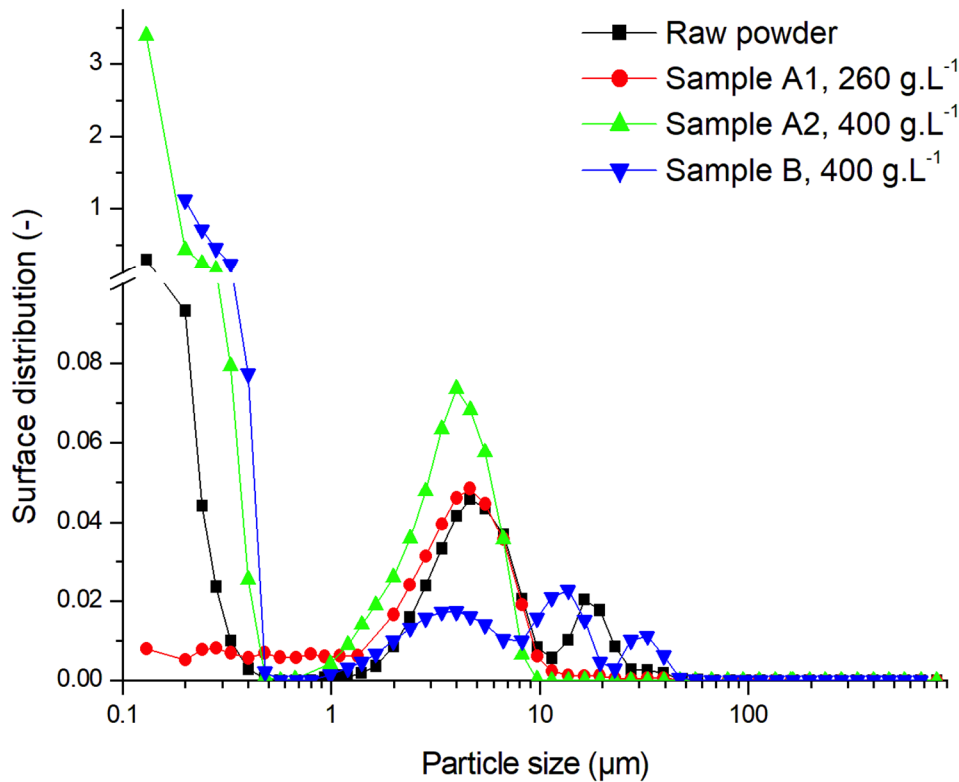
	Raw powder	Sample A1	Sample A2	Sample B
Mean diameter by wet dispersion in ethanol ( $\mu\text{m}$ ) – Mastersizer 2000 S	68	352	261	442
Mean surface diameter by sedimentation – R5 lens ( $\mu\text{m}$ ) – Helos laser diffraction sensor	207	249	132	297
Mean surface diameter by sedimentation – R3 lens ( $\mu\text{m}$ ) – Helos diffraction sensor	30	26	32	20
Mean surface diameter after dispersion in the 20 L sphere – R1+R3+R5 Helos diffraction sensor	9	4	3.5	12



Mean surface diameter after dispersion in the Scirocco device ( $\mu\text{m}$ ) – Mastersizer 2000	3.5	5.3	5	20
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297

298 To estimate the PSD during the combustion and specifically at the moment of ignition,  
 299 dispersion tests were performed with the laser diffraction sensor coupled with the  
 300 visualization 20 L sphere. Nevertheless, as seen in Table 2, each lens gives a peculiar and  
 301 useful information on the PSD, but in the case of a nanopowder, a more global view is  
 302 necessary. Unfortunately, as previously said, no apparatus currently allow such a  
 303 measurement (from nm to hundreds of micrometers) at high frequency and concentration. An  
 304 alternative solution could consist in using the same in-situ laser sensor and repeating three  
 305 times (at least) the same dust dispersion experiment with the three lenses (R1, R3 and R5). In  
 306 order to normalize the intensities of the peaks, it may be useful to insert an internal reference  
 307 (nearly monomodal powder of low and fixed concentration, with a narrow PSD which does  
 308 not interfere with the studied powder). An instance of such application for carbon-coated  
 309 silicon is given in Figure 5.



310

311 *Figure 5. Representation of the global PSD of carbon-coated silicon powders after dispersion*  
 312 *in the 20 L ( $t_v = 60$  ms) obtained by the concatenation of PSD measurements done with R1,*  
 313 *R3 and R5 (Helos – Sympatec)*

314

315 Figure 5 allows the visualization of the PSD of the nanopowders over a wide range of sizes  
 316 (from 0.1  $\mu\text{m}$  to 875  $\mu\text{m}$ ). It appears that sample B (400  $\text{g.L}^{-1}$ ) presents the biggest  
 317 agglomerates after dispersion in the sphere, with modes at 4, 15 and 35  $\mu\text{m}$  (omitting the  
 318 primary particles and nanometric agglomerates). The powders A1 and A2 show the smallest  
 319 agglomerates after dispersion, with mean surface diameters of 4 and 3.5  $\mu\text{m}$ , respectively  
 320 (Table 2). The raw Si $\Omega$ C powder, in addition to the common agglomeration mode at  
 321 approximately 4  $\mu\text{m}$ , also presents larger structures around 20  $\mu\text{m}$  (Figure 5). It should be  
 322 noticed that, unlike the other samples, no significant peak is obtained for the powder A1 at

323 particle size lower than 0.5  $\mu\text{m}$  which may seem surprising and requires additional tests with  
324 R1 lens.

325 These observations can lead to several comments: i) The measuring range being limited at  
326 0.1/0.2  $\mu\text{m}$ , the particle size distributions may be different below this value, and the presence  
327 of individualized nanoparticles is certainly more marked for the raw powder. ii) The dust  
328 injection in the 20 L sphere clearly induces the fragmentation of the biggest structures (Table  
329 2); however, the process B seems to generate more cohesive agglomerates over a wider range  
330 of particle sizes. iii) During the agglomeration process, especially with the process A, the  
331 structure of the initial agglomerates could have been weakened, reducing the cohesion  
332 strength of these initial agglomerates. During particle size measurements by wet dispersion or  
333 by sedimentation in air, their structure is not significantly modified. On the contrary, the very  
334 high shear rate induced by the dispersion in the 20 L sphere (pressurization at 20 bar) can  
335 break these agglomerates, demonstrating the existence of a threshold stress.

336

337 Table 2 and Figure 5 demonstrate the fragmentation of agglomerated structures during the  
338 injection of Si $\Omega$ C in the 20 L sphere, but quantitative indicators can be proposed. For  
339 instance, a diameter variation, defined by  $\Delta D = 100 \cdot \frac{(d_{\text{agg}} - d_{50})}{d_{\text{agg}}}$  was calculated by Sanchirico  
340 et al. (2015), who classified the powders into two classes: class 1 containing the hardest dusts  
341 ( $\Delta D < 50\%$ ) and class 2 for powders that undergo greater breakage ( $\Delta D > 50\%$ ). By  
342 considering only the smaller agglomerates (lens R3), values between 75 % and 98 % were  
343 obtained in this work, implying that all the powders are in class 2. It should be underlined that  
344 those values are slightly higher than that presented by Sanchirico et al. (2015). Nevertheless,  
345 these authors collected the dust several minutes after dust settling, which potentially allows

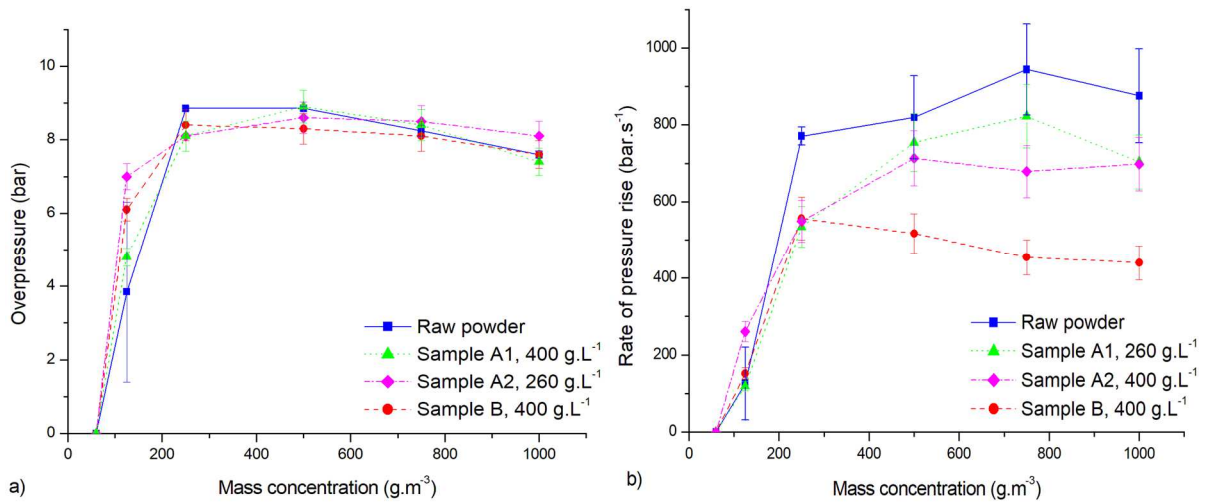
346 the re-agglomeration of the powder, whereas the measurements in this work were conducted  
347 directly within the sphere and correspond to the moment and place of ignition.

348 Dry PSD measurements were also conducted using the Mastersizer 2000 particle size analyzer  
349 equipped with a dispersion system called Scirocco (Malvern Instruments). This apparatus  
350 allows the measurement of particle sizes from 0.2 to 2000  $\mu\text{m}$ . The powder dispersion is  
351 ensured before the measurement by an air jet in a Venturi configuration (Ali et al., 2015;  
352 Bonakdar et al., 2016). Although this measurement is not conducted at the same conditions as  
353 the explosion tests, the evaluation of the particle size distribution after dispersion of the  
354 powder still provides information on the potential agglomerate breakage during the  
355 dispersion. The results, also presented in Table 2, are consistent with the SEM images, as the  
356 powder B seems the more agglomerated, with a mean surface diameter of 20  $\mu\text{m}$ , whereas the  
357 raw powder presents the smallest particles and a mean surface of 3.5  $\mu\text{m}$ . It should be noted  
358 that the shear stress induced by the dispersion in the Scirocco device is much lower than that  
359 induced during the injection in the 20 L sphere, notably due to a lower dispersion pressure (2  
360 barg). Such values tend to show that most of the initial agglomerates are broken during the  
361 dispersion in the 20 L sphere, but not necessarily in the Scirocco device, which corroborates  
362 that the agglomerates fragmentation depends on the application of a threshold stress.  
363 However, a whole characterization of the particle size distribution, from the nanometer to  
364 several hundred of micrometers and for both apparatuses, is necessary to conclude on this  
365 phenomenon.

366

367 Explosion tests were performed on each sample in the standard 20 L sphere (Figure 6). First,  
368 it should be noticed that the increase of overpressure and rate of pressure rise with the dust  
369 concentration is very fast, which tends to decrease the accuracy at the ‘transition’  
370 concentration, i.e. 125  $\text{g}\cdot\text{m}^{-3}$ . Indeed, in this transition zone, a small variation in the dust

371 concentration can greatly impact the pressure-time evolution. The maximum overpressures  
 372 obtained are similar independently of the agglomeration state of the sample, reaching around  
 373 8.8 bar (Figure 6a), which implies that the same amount of powder seems to react. From the  
 374 observation of the overpressure evolution with the concentration, agglomeration seems to  
 375 have a limited effect on the thermodynamic development of the explosion. On the other hand,  
 376 significant differences appear when regarding the maximum rate of pressure rise obtained for  
 377 each powder (Figure 6b). The raw powder leads to the most severe explosion, with a  
 378 maximum rate of pressure of  $944 \pm 118 \text{ bar}\cdot\text{s}^{-1}$ . The less agglomerated (at least theoretically)  
 379 powder i.e. SiQC A1, is slightly less severe and reaches  $822 \pm 98 \text{ bar}\cdot\text{s}^{-1}$ . Both powders are  
 380 then classified in the St2 class, with explosivity index  $K_{St}$  of respectively  $256 \pm 32 \text{ bar}\cdot\text{m}\cdot\text{s}^{-1}$   
 381 and  $223 \pm 27 \text{ bar}\cdot\text{m}\cdot\text{s}^{-1}$ . The explosions produced by the powders agglomerated with a density  
 382 of  $400 \text{ g}\cdot\text{L}^{-1}$  are less severe, reaching  $713 \pm 85 \text{ bar}\cdot\text{s}^{-1}$  for the sample A2 and only  $556 \pm 60$   
 383  $\text{bar}\cdot\text{s}^{-1}$  for the sample B, which ranks both powders as St1.



385 *Figure 6. Evolution of the a) maximum overpressure and b) maximum rate of pressure rise as*  
 386 *a function of the mass concentration of the four samples of carbon-coated silicon*

387

388 Then, it appears that only the maximum rate of pressure rise is affected by the agglomeration  
389 state of the powder. Considering spherical particles, the density of each powder and the mean  
390 surface diameter measured after dispersion in the 20 L sphere (Table 2), the reactive surface  
391 area developed after dispersion was calculated for each sample at each concentration.  
392 Contrary to the obvious correlations between the explosion severity and the reactive surface  
393 observed for starch powders (Santandrea et al., 2019a), no clear relationship could be  
394 extracted for the carbon coated silicon. Indeed, if the combustion of organic powder is limited  
395 by the pyrolysis step or by the surface reaction of the particles (Di Benedetto et al., 2010), the  
396 carbon-coated silicon does not undergo such pyrolysis step. Since the combustion does not  
397 occur in gaseous phase, the distance between the particles is of great importance. Indeed,  
398 increasing the concentration in the same volume reduces this distance, possibly leading to a  
399 transition between the combustion of the individual particle (small concentration) to a group  
400 combustion at higher concentrations. On the contrary, promoting the agglomeration tends  
401 both to increase the average distance between the combustible structures and to decrease the  
402 surface concentration.

403 After each explosion test, the evolution of the pressure was recorded during the cooling phase.  
404 Since the combustion of silicon consumes oxygen and produces silica, the final pressure in the  
405 sphere is lower than the atmospheric pressure, even if the combustion of the carbon layer  
406 leads to the production of, among others, carbon monoxide and carbon dioxide. The stabilized  
407 pressure after explosion can provide information on the combustion gases production and  
408 oxygen consumption (Figure 7).

409 To evaluate the combustion mechanisms, two hypotheses were investigated. First, successive  
410 reactions were considered: the carbon C reacts first, followed by the silicon Si (Figure 7). The  
411 second hypothesis consists of simultaneous reactions of the carbon and the silicon, the oxygen  
412 consumption by each reaction being proportional to the molar proportions of C and Si. The

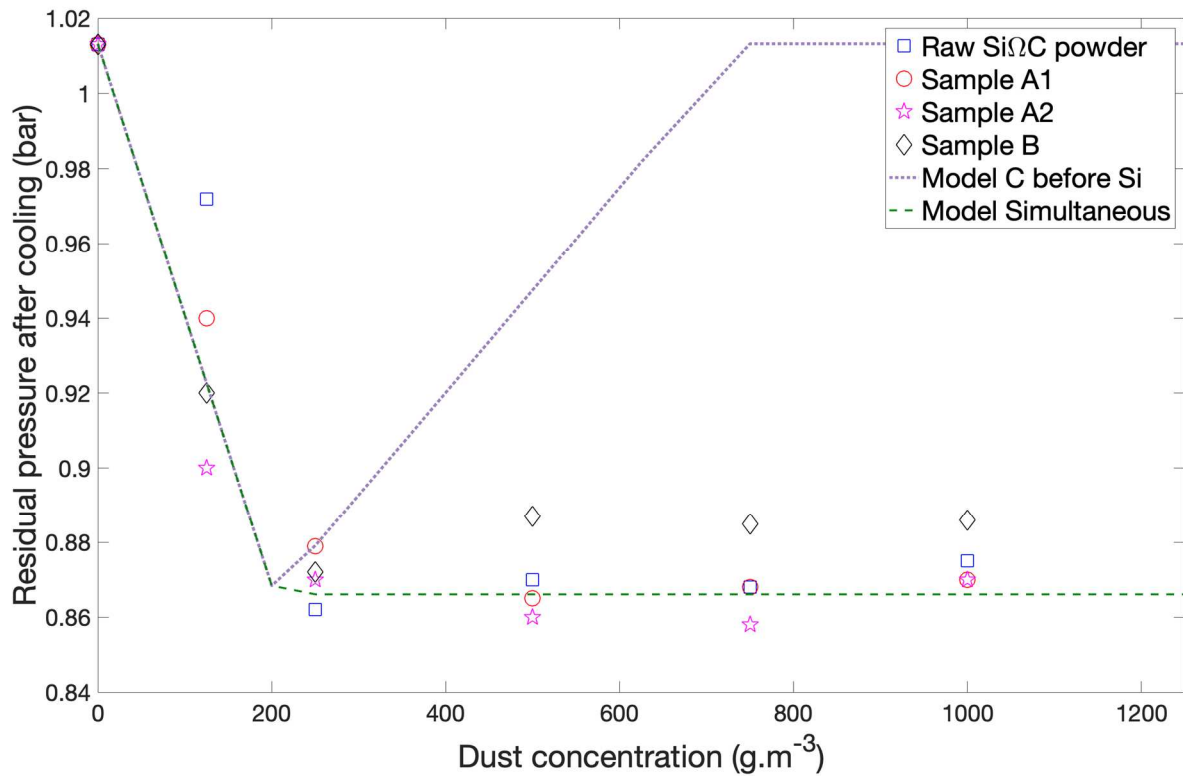
413 considered reactions are the formation of carbon dioxide from the carbon and oxygen, and the  
414 reaction of silicon with oxygen to form silica ( $\text{SiO}_2$ ). Assuming that the particle diameter is  
415 40 nm with 2 nm of carbon coating, a molar proportion of 27.4% of carbon and 72.6% of  
416 silicon was considered. In light of the similar measured overpressures, it was previously  
417 established that the reaction products were not significantly affected by the agglomeration  
418 state. For a given mass concentration, the same amount of powder is then assumed to be  
419 consumed for each sample and the reaction of nitrogen with silicon was neglected.

420 It appears in Figure 7 that the experimental residual pressure seems to follow the same  
421 evolution than in the case of simultaneous reactions of carbon and silicon, even if the carbon  
422 layer is the first compound theoretically in contact with oxygen. Such behavior can notably be  
423 explained by the combustion of silicon in vapor phase, which is confirmed by considering  
424 Glassman's criterion: the boiling point of Si ( $2355^\circ\text{C}$ ) is much lower than the volatilization  
425 temperature of its oxide ( $2950^\circ\text{C}$  for silica) or of its carbon layer (more than  $4000^\circ\text{C}$ ). As a  
426 consequence, after ignition, i.e. at high temperatures, the carbon coating does not prevent or  
427 hinder the combustion of the silicon. If the second model seems to represent satisfactorily the  
428 experimental data, a concentration shift is visible for the low dust concentrations. Indeed, the  
429 value experimentally measured at  $125 \text{ g}\cdot\text{m}^{-3}$  for the raw powder, i.e. around 0.97 bara,  
430 corresponds to the value calculated at  $60 \text{ g}\cdot\text{m}^{-3}$  by both models. As a lower pressure would  
431 have been theoretically expected, this shift may be due to oxygen adsorbed onto the particles  
432 surface, reacting more easily with the carbon layer and generating an extra amount of carbon  
433 dioxide. Moreover, Figure 7 shows that the powders agglomerated by the process A, i.e.  
434 powders A1 and A2, follows globally the same evolution than the raw powder with a final  
435 pressure stabilized around 0.87 bara. On the contrary, the final pressure obtained after  
436 combustion of powder B stabilized around 0.89 bara, which confirms the specificity of this  
437 powder due to the agglomeration process. Due to the PSD of sample B being greater than

438 those of the other Si $\Omega$ C powders, the heating rate of such structures is probably slower, which  
439 may modify the reaction mechanism, promoting the effect of the carbon layer and thus  
440 leading to a lowest explosion severity.

441





443

444 *Figure 7. Evolution of the stabilized pressure after explosion in the 20 L sphere with the mass*  
 445 *concentration: experimental data for the different samples of carbon-coated silicon and*  
 446 *models for the raw powder*

447

448 In addition to the experiments conducted in the 20 L sphere, the minimum ignition energy  
 449 (MIE) of the different powders was determined in the modified Hartmann tube according to  
 450 ISO/IEC 80079-20-2 (2016). While the MIE of the raw powder was evaluated at 5 mJ, no  
 451 ignition was observed at 1J for the three other samples when varying the dust quantity from  
 452 0.6 to 1.8 g and the ignition delay time from 90 ms to 150 ms. Although some glowing  
 453 particles were observed when testing the sample A1, no flame was obtained. The  
 454 agglomeration then appears to significantly reduce the ignition sensitivity of SiΩC powders,  
 455 turning very sensitive raw powders to samples almost insensitive to electrostatic ignition

456 (Janès et al., 2008). By examining the evolution of the PSD in the tube with the R5 lens, it  
457 appears that mean surface diameters of 82 and 252  $\mu\text{m}$  were obtained for raw Si $\Omega$ C sample  
458 and sample B, respectively; which is consistent with MIE results.

459 Nevertheless, as shown in Figure 6, the lowest concentration for which an explosion was  
460 obtained at 10 kJ is the same for all the samples, i.e. 125  $\text{g}\cdot\text{m}^{-3}$ . Since these tests were  
461 performed at 10 kJ, additional tests should be conducted with the standard ignition energy to  
462 apply to determine the LEL, i.e. 2 kJ (EN 14034-3, 2006). Still, the LEL does not seem  
463 significantly modified by the agglomeration state of the powder, which can notably be  
464 explained by the strong fragmentation of the agglomerates during the dispersion in the 20 L  
465 sphere. Indeed, the agglomerates submitted to the high shear stress of the injection device in  
466 the 20 L sphere tend to break more easily than the agglomerates lifted by the air pulse in the  
467 modified Hartmann tube, which give them ignition properties similar than the raw powder.  
468 This point will be developed in the next section.

469

### 470 **3.3. Some theoretical clues on relating the dispersion process and the particle size** 471 **distribution of nanopowders clouds**

472 Inertia, rotary and turbulent stresses are the three main stresses exerted by the fluid flow on  
473 the particles surface. It seems important to consider them as their evolution can explain the  
474 deagglomeration phenomenon and thus, the changes in the ignition sensitivity and explosivity  
475 previously described.

476 The inertia stress is due to agglomerates shifts and accelerations in the flow direction. It can  
477 be defined by equation 1 (Deng et al., 2016; Weiler et al., 2010):

$$478 \sigma_l = \frac{\rho}{8} \cdot \left[ 1 - \frac{\arccos\left(\frac{d_p}{d_{agg} - d_p}\right)}{180^\circ} \right] \cdot (u_p - u)^2 \cdot \left[ \frac{24}{\text{Re}_{rel}} + \frac{24}{\text{Re}_{rel}^{0.5}} + 0.4 \right] \quad (1)$$

479 where  $\rho$  is the fluid density;  $d_p$ , its primary diameter;  $d_{agg}$ , the diameter of the agglomerate;  $u$ ,  
 480 the flow velocity;  $u_p$ , the agglomerate velocity and  $Re_{rel}$ , the Reynolds number calculated  
 481 from the slip velocity  $u_s$  between the agglomerates and the gas:

$$482 \quad u_s = (u_p - u) = d_p^2 \frac{(\rho_p - \rho)}{18\mu} \cdot a \quad (2)$$

483 where  $\rho_p$  is the particle density;  $\mu$  the fluid dynamic viscosity and  $a$  the acceleration term.

484 The rotary stress  $\sigma_r$  is generated in zones with high velocity gradient ( $du/dz$ ). Weiler et al.  
 485 (2010) defines it as follows:

$$486 \quad \sigma_r = \frac{\rho_p}{40} \cdot (d_{agg})^2 \cdot \left(\frac{du}{dz}\right)^2 \quad (3)$$

487 Finally, turbulent stresses  $\sigma_T$  caused by the vortices present in the velocity field can also lead  
 488 to agglomerate breakage both due to impaction between solid structures and to shear stresses.  
 489 They depend on the ratio between the Kolmogorov scale and the agglomerate size and are  
 490 expressed as a function of fitting parameters  $a_i$  given by Weiler et al. (2010):

$$491 \quad \sigma_T = a_1 \rho \cdot \left[ \frac{v_{kin}^{(3a_2-1)}}{l_D^{4 \cdot a_2}} \right]^{a_3} \cdot (d_{agg})^{a_4} \quad (4)$$

492 where  $v_{kin}$  is the kinematic viscosity of the fluid and  $l_D$ , the Kolmogorov scale of the flow.  
 493 The latter parameter can be found for the 20 L sphere at various dispersion times, especially at  
 494 10, 50 and 100 ms for Kolmogorov scale (Dahoe et al., 2001; Ogle, 2016). Due to  $l_D$  values  
 495 being unavailable for dust dispersions in the Hartmann tube, they have been evaluated by the  
 496 following equation:

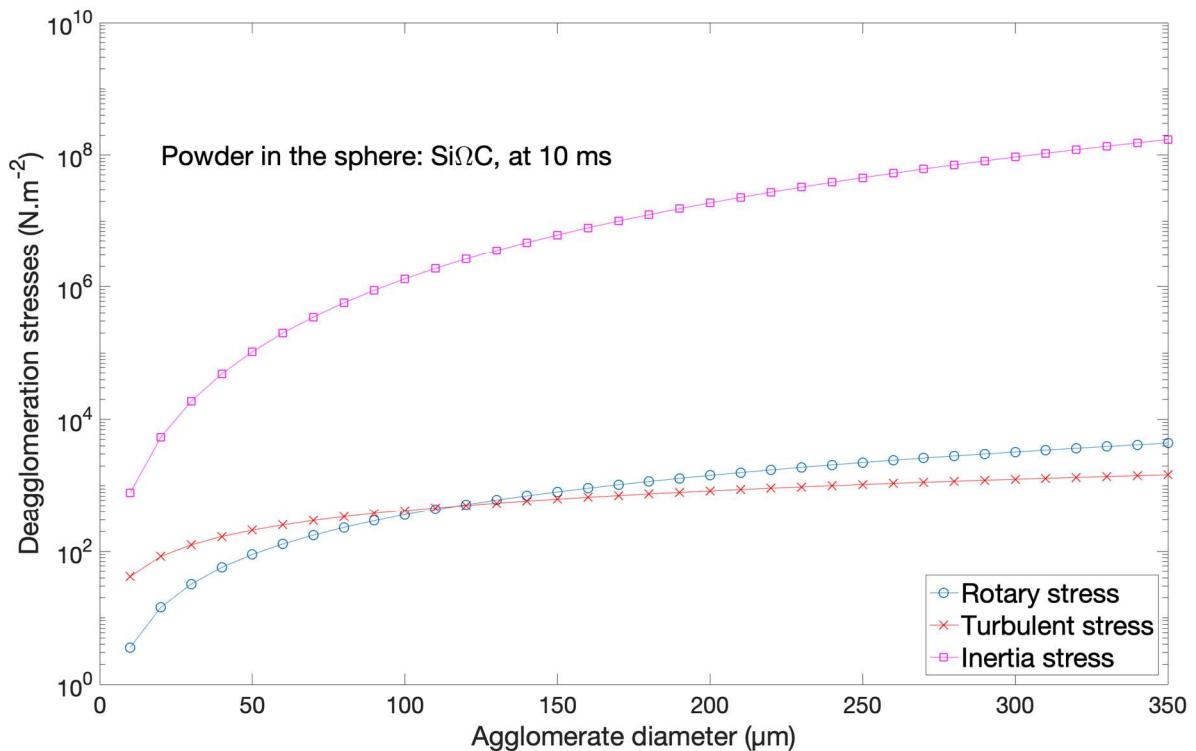
$$497 \quad l_D = \left( \frac{v_{kin}^3}{\varepsilon} \right)^{0.25} \quad (5)$$

498 where  $\varepsilon$  is the turbulence dissipation rate which can be estimated as follows by considering  
 499 only the vertical direction  $z$ :

$$500 \quad \varepsilon = 15 \cdot v_{kin} \cdot \left(\frac{du_z}{dz}\right)^2 \quad (6)$$

501 This assumption made on a privileged direction of the particle flow is confirmed by previous  
502 Particle Image Velocimetry (PIV) experiments (Cuervo, 2015). Particles and agglomerates  
503 velocities, velocity gradients and acceleration were determined both for the sphere and the  
504 tube by CFD simulations and were validated experimentally by PIV measurements (Cuervo,  
505 2015; Murillo, 2016).

506 Figure 8 shows that the inertia stress is by far the greatest stress which applies on the  
507 agglomerates, significantly increasing when  $d_{agg}$  increases. The predominance of drag forces  
508 on other fragmentation mechanisms is confirmed by other authors (Breuer and Khalifa, 2019)  
509 and is illustrated by a Stokes number greater than 90 for a dust dispersion of Si $\Omega$ C in the  
510 sphere, corresponding to large relative velocities between the agglomerates and the flow.  
511 However, other mechanisms such as the impaction of large particles on the upper and lower  
512 plates of the rebound nozzle (Kalejaiye et al., 2010) should not be neglected. It should be  
513 stressed that the calculations were performed for the first 10 ms of the dust dispersion as high  
514 acceleration rates specifically occur on short time scales (Murillo, 2016; Weiler et al., 2010).



515

516 *Figure 8. Evaluation of the intensities of the rotary, turbulent and inertia stresses in the 20 L*  
517 *sphere for Si $\Omega$ C agglomerates*

518

519 Figure 9 represents the evolution of the global stress (sum of inertia, rotary and turbulent  
520 stresses) exerted on Si $\Omega$ C agglomerates both in the 20 L sphere and in the modified Hartmann  
521 tube. At first, it can be noticed that, whatever the dispersion time, the stress which applied on  
522 the agglomerates in the 20 L sphere is always greater than in the modified Hartmann tube. It  
523 can explain why the agglomerated samples of Si $\Omega$ C showed a minimum ignition energy  
524 greater than 1 J as they cannot be fragmented during their dispersion in the tube. On the  
525 contrary, Figure 5 demonstrates that the Si $\Omega$ C samples can be deagglomerated, at least  
526 partially, by their injection in the 20 L vessel. As previously said, Figure 9 also confirms that  
527 the fragmentation occurs during the first moments of the dust dispersion, when the  
528 acceleration and the slip velocity are the greatest. Three models were used to assess the  
529 agglomerate strength  $\sigma_{agg}$  and compare it to the deagglomeration stress (Kendall, 1988;  
530 Rumpf, 1962; Weiler et al., 2010). Rumpf model assumes that the agglomerate is separated in  
531 two parts by a transverse section, without taking into account the actual structure of the  
532 agglomerate. It implies that the breakage occurs through a simultaneous rupture of all the  
533 bonds along the fracture plan:

534 
$$\sigma_{agg,R} = \frac{9}{8} \cdot \left( \frac{1-\epsilon_p}{\epsilon_p} \right) \cdot \frac{F_c}{d_p^2} \quad (7)$$

535 where  $\epsilon_p$  is the porosity of the agglomerate and  $F_c$ , the cohesion force, limited here to the Van  
536 der Waals forces  $F_{vdw}$ . With regard to the work presented by Deng et al. (2016), the porosity  
537 of Si $\Omega$ C agglomerates was set at 0.8 and  $F_{vdw}$  was estimated through the following  
538 relationship valid for 2 identical spheres:

539 
$$|F_{vdw}| = \frac{H \cdot d_p}{24 \cdot h_p^2} \quad (8)$$

540 where  $H$  is the Hamaker constant and  $h_p$  is the cutoff of separation, set at 0.165 nm (Deng et  
 541 al., 2016). As Si is coated by a carbon layer, an average Hamaker constant  $H$  of  $2.5 \cdot 10^{-19}$  J  
 542 was chosen, based on the work of Dagastine et al. (2002), indicating that this value is valid for  
 543 graphite particles in air.

544 Kendall (1988) claimed that a more gradual rupture occurs and that the real cohesion strength  
 545 of an agglomerate is overestimated by Rumpf model. As a consequence, the following  
 546 relation was proposed for the cohesion strength:

$$547 \quad \sigma_{\text{agg,K}} = 15.6 \frac{(1-\varepsilon_p)^4}{d_p} \cdot \frac{H}{12 \cdot \pi \cdot h_p^2} \quad (9)$$

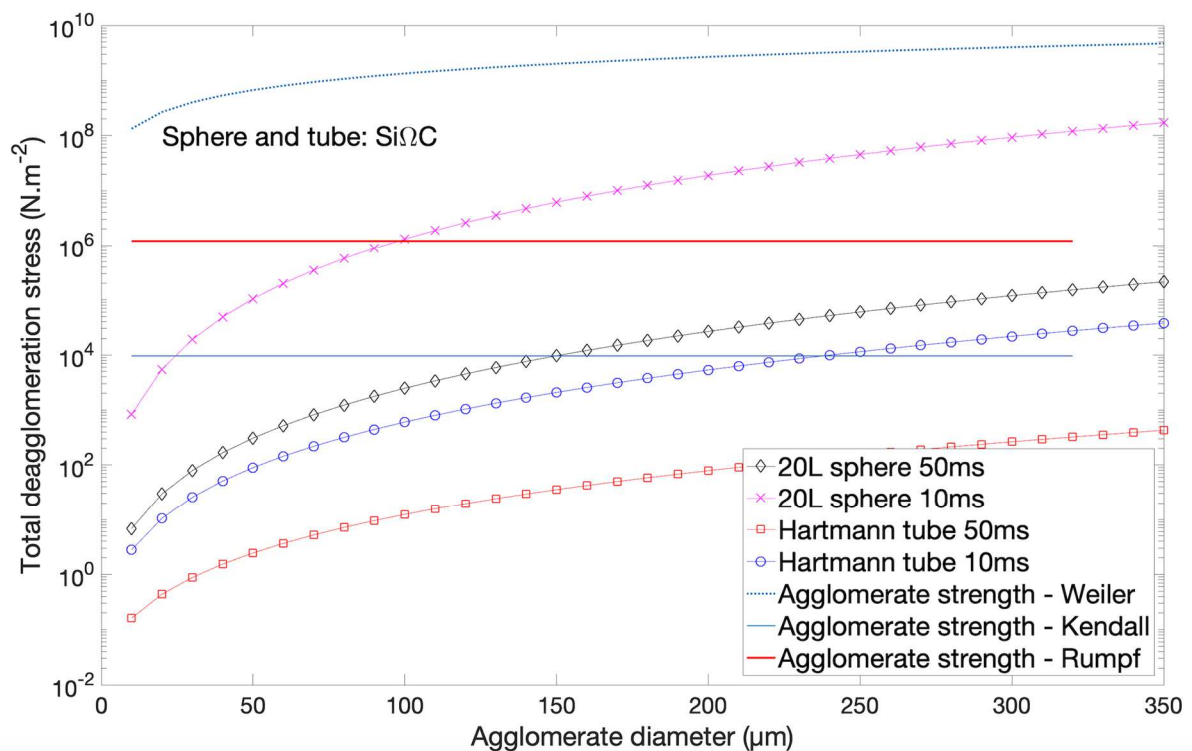
548 Finally, Weiler et al. (2010) proposed an alternative model assuming a total breakage of the  
 549 agglomerates, by considering all the contacts between the particles. Obviously, the cohesion  
 550 strength of such structure is greater than those obtained by the previous models:

$$551 \quad \sigma_{\text{agg,W}} = \frac{(1-\varepsilon_p)}{\varepsilon_p} \cdot \frac{F_c}{2 \cdot d_{\text{agg}}^2} \cdot \left[ \left( \frac{d_{\text{agg}}}{d_p} \right)^3 - \frac{4 \cdot \left( 1 - \frac{\text{Arccos}(b)}{180^\circ} \right)}{b^2} \right] \quad (10)$$

$$552 \quad \text{with } b = \frac{d_p}{d_{\text{agg}} - d_p}.$$

553 In Figure 9, it should be noticed that, even in the 20 L sphere, the stresses are not sufficient to  
 554 break every Van der Waals bond of Si $\Omega$ C agglomerates, which is notably confirmed by  
 555 Figure 5 in which microstructures are still clearly visible. By applying Rumpf model, it seems  
 556 that only the structures larger than 90  $\mu\text{m}$  will be broken during a dust injection in the 20 L  
 557 sphere, whereas the limit defined by Kendall model is close to 25  $\mu\text{m}$ . The latter value is  
 558 consistent with the measurements performed with the laser sensor (Figure 5), but no larger  
 559 agglomerates are visible after dust dispersion in the sphere, which tends to support Kendall's  
 560 approach. These conclusions are also in accordance with the Monte Carlo simulations  
 561 developed by Deng et al. (2016) on 50 nm primary particles. Indeed, they obtained cohesion  
 562 strengths of  $2.8 \cdot 10^4$  Pa and  $1.7 \cdot 10^5$  Pa for Kendall and Rumpf models, respectively.

563 Nevertheless, it should be underlined that the cohesion strengths presented in our study should  
 564 be considered as orders of magnitude as they are greatly dependent on variables which are  
 565 difficult to quantify, such as Hamaker constant and the cutoff of separation. The same  
 566 approach can be applied to nanocellulose dispersion, with similar conclusions. However, it  
 567 will not be detailed here as the fibrous nature of the primary particles makes even more  
 568 difficult to define their structural properties.



569  
 570 *Figure 9. Evaluation of the intensities of the deagglomeration stresses in the 20 L sphere and*  
 571 *in the modified Hartmann tube for SiΩC agglomerates. Comparison with cohesion strengths*  
 572 *models.*

573

#### 574 4. Conclusions

575 The ignition and explosion characteristics of nanocellulose and silicon coated by carbon were  
 576 studied with regard to their agglomeration state, through sieving and mechanical  
 577 agglomeration. It mainly appears that the agglomerates, naturally or intentionally generated,

578 can be broken if a sufficient stress is applied during their dispersion. The threshold stress can  
579 be roughly estimated using models allowing for the determination of the cohesion strength of  
580 agglomerates (Kendall model for instance).

581 When applying rather low shear stresses, as for instance during the dust dispersion in a  
582 modified Hartmann tube, only the largest structures can be broken, and the minimum ignition  
583 energy is greatly modified by the presence of big agglomerates. However, if a greater stress is  
584 applied, such as during the dust dispersion in the 20 L sphere, most of the micrometric  
585 agglomerates can be broken and the explosion severity is only slightly affected by the  
586 presence of brittle agglomerates. But when harder structures are formed, as it is the case for  
587 Si $\Omega$ C produced by the process B, the maximum rate of pressure rise is significantly decreased  
588 and the combustion mechanisms can also be altered.

589 Such modification of the explosion risk may be considered as a direct application of the  
590 moderation principle of inherent safety. By intentionally agglomerating nanopowders, their  
591 ignition sensitivity, but also to a lesser extent, their explosion severity, can be greatly reduced.  
592 Nevertheless, it should be underlined that such assertion mainly depends on the stress which  
593 will be applied to disperse the powders, during explosion tests, but above all, during an  
594 accident. It should be added that a similar approach can be applied to micron-size particles or  
595 ultrafine particles.

596 The specificity of the nanopowders lies in their high specific surface area. But a dust cloud of  
597 nanopowders (at concentrations greater than the minimum explosive concentration) is always  
598 highly polydispersed and the whole particle size distribution has to be considered. Each part  
599 of the cloud can 'play its own role': the primary nanoparticles influences greatly the ignition  
600 step, the primary aggregates or agglomerates of a few micrometers play a significant role in  
601 the radiative heat transfer and the bigger micrometric structures can be both considered as



602 flame stretching agents and ‘fuel reserves’. But if the cohesion strength of the agglomerates is  
603 too high with regard to the dispersion stress, the explosion risk will be lowered.

604

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