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Solvothermal synthesis, structure and magnetic properties of heterometallic coordination polymers based on a phenolato-oxamato co-bidentate-tridentate ligand

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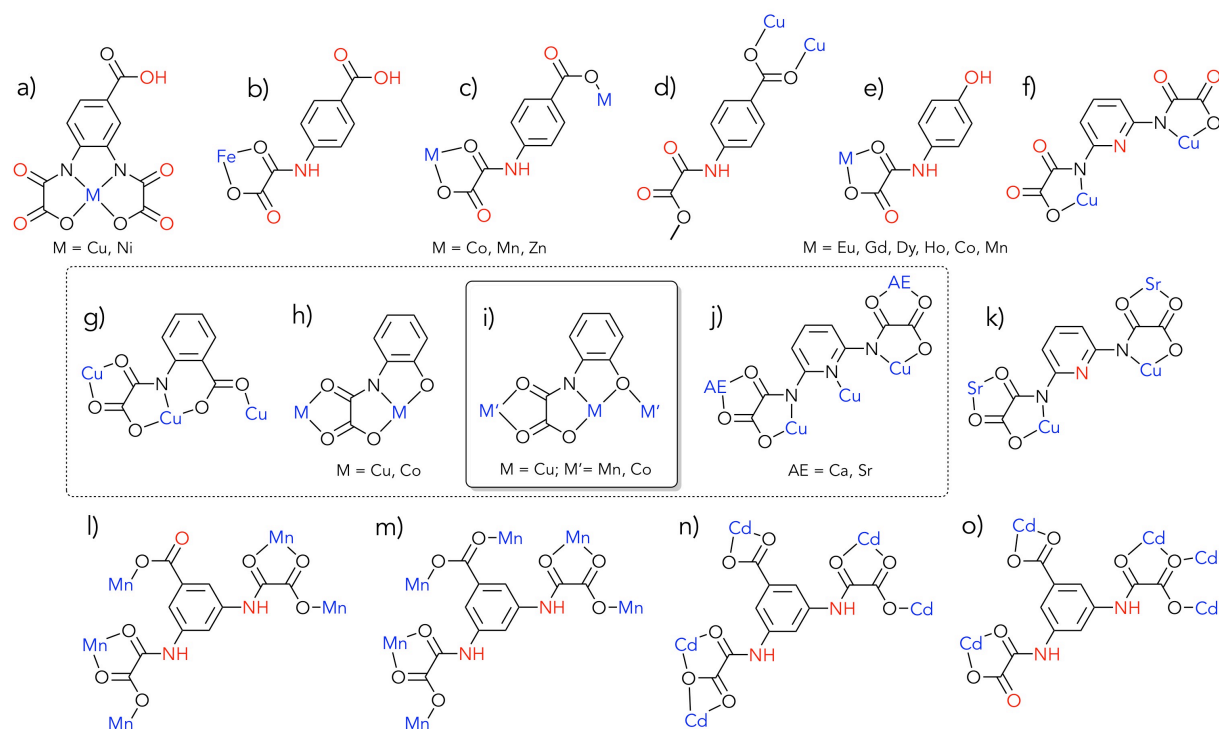
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The use of solvothermal conditions has successfully led to the preparation of heterometallic 1D coordination polymers from a co-bidentate-tridentate phenolato-oxamato ligand. The reaction of the *N*-(2-hydroxyphenyl)oxamic acid (ohpma) with acetate salts of transition metal ions at 80°C has yielded the heterobimetallic [Cu(ohpma)M(OAc)(DMF)₂] (M = Co (**1**); Mn (**2**)) and the heterotrimetallic [Cu(ohpma)Co_{0.57}Mn_{0.43}(OAc)(DMF)₂] (**3**) chain compounds. Single-crystal and powder diffraction studies show that the polymers are isostructural. Magnetic studies suggest the existence of an inter-chain two-dimensional antiferromagnetic interaction taking place in compounds **1-3**.

Introduction

N-substituted oxamate ligands provide a rich chemistry and prove to be efficient tools in coordination chemistry for the preparation of molecular materials.¹⁻³ Looking at aromatic derivatives, rational design has motivated the preparation of numerous ligands that bear bulky, chiral, electro- and photo-active substituents in order to perfect the control over the shape, the dimensionality and the properties of the final edifices.⁴⁻⁶ However, the investigation of *multi*-polydentate ligands bearing an additional coordinating group to the oxamate one has been far less looked into. The phenylene-bisoxamate scaffold bearing an additional carboxylato group has been reported to yield homometallic compounds, yet their study has been limited to the metalloligand (Scheme 1a) or revealed synthetic issues in obtaining a fully deprotonated oxamate function (Scheme 1, l-o).⁷⁻¹⁰ In either case, a “whole” coordination of the ligand where all the N/O donor atoms participate was not achieved. A pyridine-based bis-oxamate ligand has permitted the observation of a full coordination mode in the multi-polydentate aromatic oxamate-based ligands family (Scheme 1j),¹¹ yet it has also yielded partially binding forms where the pyridine group is free (Scheme 1f, 1k).^{11,12} Similar issues have been observed with the para-substituted phenyloxamate ligands, with either carboxylato or hydroxido additional groups where the oxamate and the additional coordinating groups are only partially or not at all binding (Scheme 1, b-e).¹³⁻¹⁶ On the other hand, the ortho-substituted carboxylato-phenyloxamate ligand has been reported entirely coordinating in homometallic copper-based polymers (Scheme 1g).¹⁷ These observations led us to

recently investigate the reactivity of the ortho-substituted N-(2-hydroxyphenyl)oxamic acid (H_3ohpma).¹⁸ We have obtained two homometallic compounds also displaying a whole coordination mode with copper or cobalt ions (Scheme 1h). Nevertheless, we failed at preparing heterometallic systems, and besides the compounds made of Cu(II) and alkaline earth cations,¹¹ none of the compounds reported so far with phenyloxamate-based ligands bearing additional coordinating groups are heterometallic systems. The later are of importance in terms of magnetic properties, and synthetically, one of the key features offered by oxamate chemistry. Indeed, the dissymmetry the bis-bidentate oxamate group facilitates the preparation of heterometallic pairs. Considering that our initial attempts in standard conditions failed in the crystallisation of the heterometallic compounds and not in reacting the ligand, we have envisaged the use of solvothermal conditions. Giving its affinity for the $\kappa N, \kappa O$ chelating site of the oxamate ligands,¹⁹ we assumed that using copper(II) as one of the metal ions would circumvent the one-pot criterion of the solvothermal method. Indeed, a multi-step synthesis is the main strategy on the bench when preparing heterometallic coordination compounds with oxamate ligands.²



Scheme 1. Drawings of the known coordination modes for oxamate-based multi-polydentate ligands bearing additional coordinating groups.

Our approach was successful, and we present here the solvothermal reaction of the H_3ohpma ligand with transition metal ions. It has yielded a novel 1D polymer in the form of three heterometallic analogues, $[Cu(ohpma)Co(OAc)(DMF)_2]$ (**1**), $[Cu(ohpma)Mn(OAc)(DMF)_2]$ (**2**) and $[Cu(ohpma)Co_{0.57}Mn_{0.43}(OAc)(DMF)_2]$ (**3**). All the compounds have been characterised by single-crystal and powder X-ray diffraction, and investigated for their magnetic properties.

Results and discussion

Compounds **1** and **2** are isostructural and consist of a neutral hetero-bimetallic oxamato-, phenolato- and acetato-bridged chain (Figure 1) that crystallizes in the $P2_1/c$ monoclinic space group. The chain is made of $[Cu(ohpma)]^+$ complexes where the Cu^{II} ion is coordinated to one

nitrogen atom (N1) and one oxygen atom (O1) of the oxamate group, and the oxygen atom (O4) of the phenolato group. The complexes then bridge the additional M^{II} ions (Co^{II} in **1**, Mn^{II} in **2**) in one direction via the remaining chelating carbonyl groups of the oxamate function (O2, O3), and in the other via the m-oxo atom of the ligand's phenolato group as well as an additional m-acetato bridge (O5, O6). This coordination mode for the (ohpma)³⁻ ligand (μ -oxamato- κ^2 O,O': κ N, κ O''- μ -phenolato, Scheme 1i) differs from the one we have previously observed for homometallic compounds prepared in bench conditions (μ -oxamato- κ^2 O,O': κ N, κ O''-phenolato- κ O, Scheme 1h). Here, we observe a bridging phenolato oxygen atom instead of a terminal one.

The copper ions are four-coordinate in a square-planar environment with Cu-O distances (1.956 Å [1.902(1)-2.006(1)]), and Cu-N bond length (1.902(2) and 1.905(1) Å) that are comparable to those found in the homometallic 1D polymer, and are indicative of a constrained environment.¹⁸ The O/N-Cu-O/N angles range from 83 to 102°, and the dihedral angle between the N1-Cu-O1 and the O4-Cu-O5 planes are of 13 and 12°, for **1** and **2** respectively. The M^{II} ions are six-coordinate, with two DMF molecules completing the coordination sphere. They have fairly regular octahedral environments, with homogeneous M-O distances ranging from 2.061(1) to 2.133(2) Å for **1**, and from 2.131(2) to 2.206(2) Å for **2**. These distances are similar to previously reported oxamate-based 1D polymers with *cis*-coordinated Co^{II} or Mn^{II} ions²⁰⁻²². The O-Co-O angles vary from 79 to 98°, and deviate on average by 4° from orthogonality. For Mn^{II} ions, these angles range from 76 to 100°, with an average deviation of 5°.

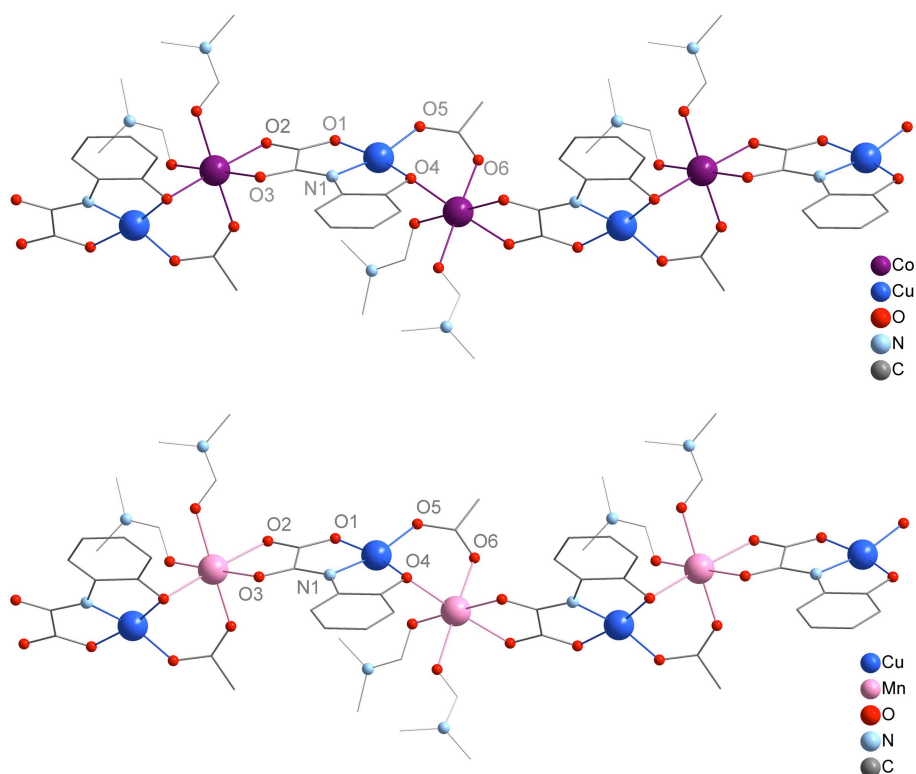


Figure 1. Representation of the heterometallic chains **1** and **2**, with atom labels. H atoms have been omitted for clarity.

The chains are parallel to each other and aligned along the crystallographic *c* axis (see supplementary information). The shortest interchain metal-metal distances were found at 7.12 Å (Cu-Co in **1**) and 7.15 Å (Cu-Mn in **2**). Furthermore, the chains form sheets perpendicular to the

crystallographic *a* axis (Figure 2). These layers of chains are alternatively separated by aromatic rings and DMF molecules and the resulting distances between mean planes of magnetic ions are 8.46 Å (8.43 in **2**) and 6.69 Å (6.58 in **2**).

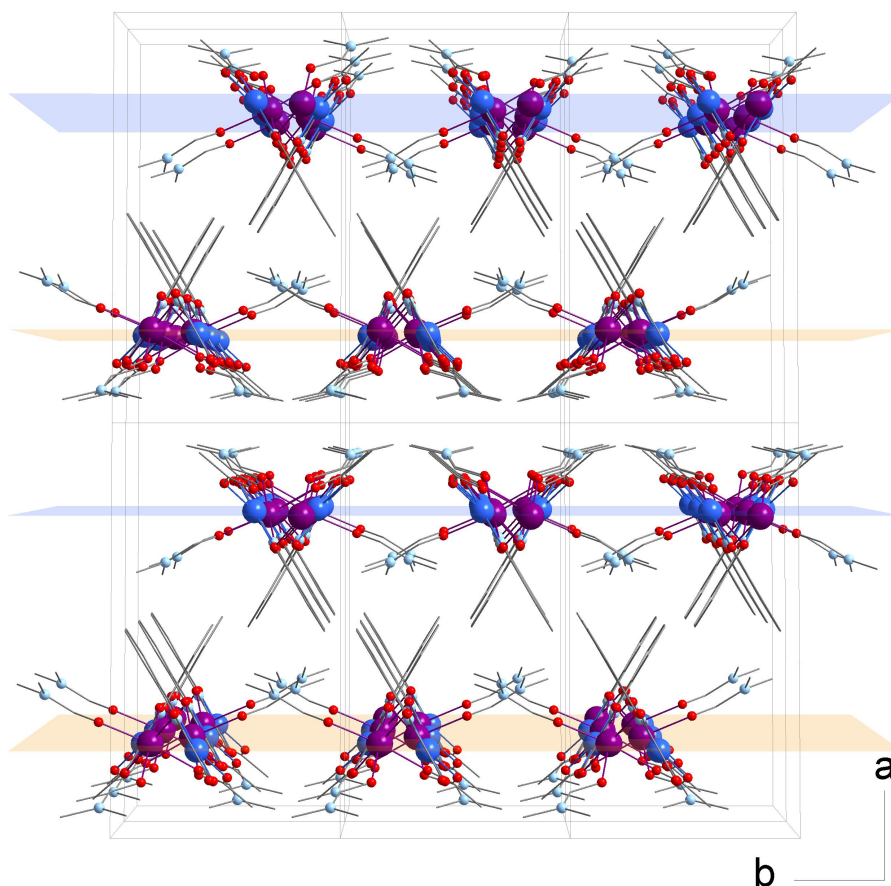


Figure 2. Representation of the layers of chains in **1**. H atoms have been omitted for clarity.

To further investigate the magnetic properties, we have introduced manganese ions in the synthesis of the $[\text{Cu}(\text{ohpma})\text{Co}(\text{OAc})(\text{DMF})_2]$ chain. This has led us to the following heterotrimetallic compound: $[\text{Cu}(\text{ohpma})\text{Co}_{0.57}\text{Mn}_{0.43}(\text{OAc})(\text{DMF})_2]$ (**3**). Single-crystal data (see Table 2 in the Experimental section) and Le Bail refinement of the powder X-ray patterns (Table 2 and Figure 3) performed on the bulks of compounds **1-3** confirm that all three chains are isostructural. The differences that are observed between the experimental X-ray powder diffraction diagrams reflect the small differences observed with the single-crystal crystallographic parameters. Furthermore, the Le Bail refinements indicate mainly a slight increase of the unit cell parameter *a* at room temperature, which could be consistent with the packing of the chains in the solid. The metal content in the formula of compound **3** was taken from ICP-AES results and differ from the single-crystal X-Ray data. In the latter, and giving the quality of the data set, site occupancies for the Co and Mn atoms were refined and gave on the measured crystal $\{\text{Co}_{0.8}\text{Mn}_{0.2}\}$ for **3**. A second set of ICP-AES titration confirmed the ratio given in the formula and was $\{\text{Co}_{0.56}\text{Mn}_{0.44}\}$.

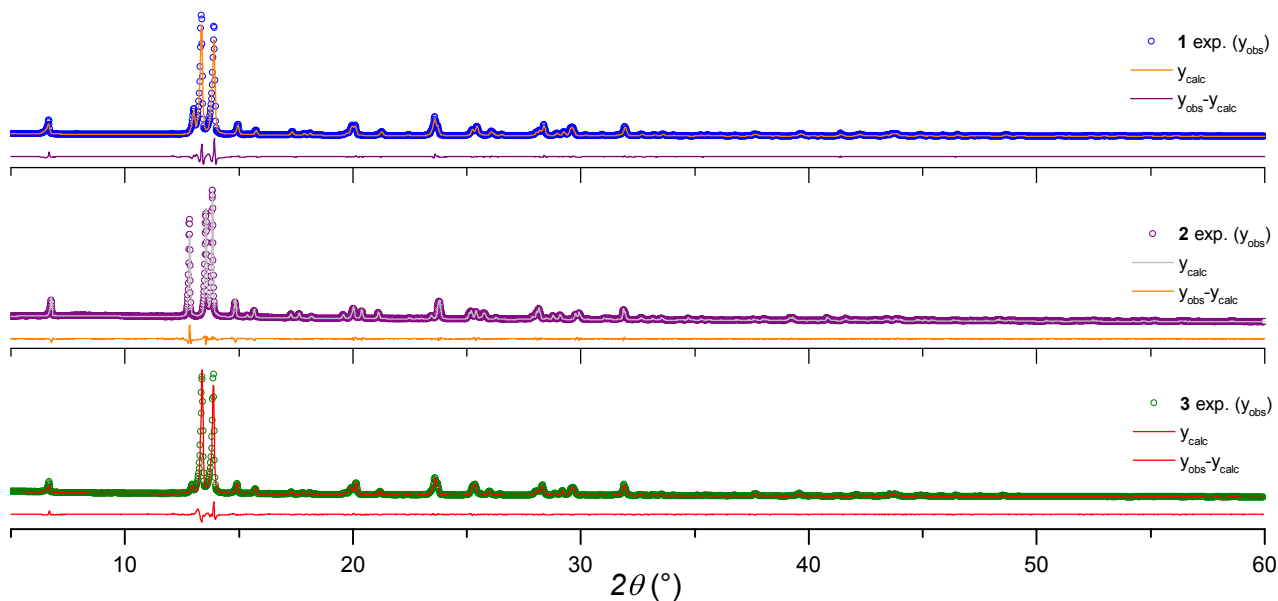


Figure 3. Experimental and refined powder patterns for compounds **1-3**.

The DC magnetic properties of compounds **1-3** in the form of $\chi_M T$ versus T plots are shown in Figure 3. The value of $\chi_M T$ for compound **1** is within the theoretical range, assuming non-interacting ions (see Table 1). In the presence of Mn^{II} however, these values are below the theoretical limit, pointing at the existence of a strong antiferromagnetic interaction in **2** and **3**, or at least, stronger than in **1**.

	Exp.	Theo.	
1	3.41	$3.11 < \chi_M T < 3.81$	$\text{Cu}^{\text{II}}: S_{\text{Cu}} = 1/2; g_{\text{Cu}} = 2.1; \chi_M T = 0.41$
2	4.5	4.78	$\text{Co}^{\text{II}}: {}^{23} 2.7 < \chi_M T < 3.4$
3	3.81	$3.83 < \chi_M T < 4.23$	$\text{Mn}^{\text{II}}: S_{\text{Mn}} = 5/2; g_{\text{Mn}} = 2.0; \chi_M T = 4.37$

Table 1. Experimental and theoretical $\chi_M T$ values ($\text{cm}^3\text{Kmol}^{-1}$) for **1-3**.

The existence of antiferromagnetic interactions is supported by the temperature dependence of the $\chi_M T$ products, which, in all cases, continuously decrease from 300 to 60 K (Figure 3). In the case of **2**, $\chi_M T$ reaches a minimum at 60 K, with a value of $3.88 \text{ cm}^3\text{Kmol}^{-1}$, before increasing sharply up to $16.85 \text{ cm}^3\text{Kmol}^{-1}$ at 2 K. For **3**, the decrease of $\chi_M T$ continues from 60 K to 18 K where a plateau at $2.05 \text{ cm}^3\text{Kmol}^{-1}$ is reached. The $\chi_M T$ product then slightly increases to $2.20 \text{ cm}^3\text{Kmol}^{-1}$ at 4.5 K before reaching 2.11 at 2 K. The $\chi_M T$ product of **1** keeps decreasing below 60 K, down to $0 \text{ cm}^3\text{Kmol}^{-1}$ at 2 K. A small field-dependent contribution was observed at 15 K in the heterobimetallic chain **1**. This is confirmed by the presence of two weak anomalies in the zero-field-cooled and field-cooled (ZFC/FC) magnetisation curves at 15 K and 5.5 K that are due to the presence of small amounts of magnetic phases showing a ferromagnetic ordering. This behaviour is attributed to trace amounts of a non stoichiometric cobalt hydrate hydroxide impurity, that presents a magnetic ordering in this temperature range and are unfortunately often observed

when performing low field measurements on cobalt(II)-based compounds (see supplementary information).²⁴

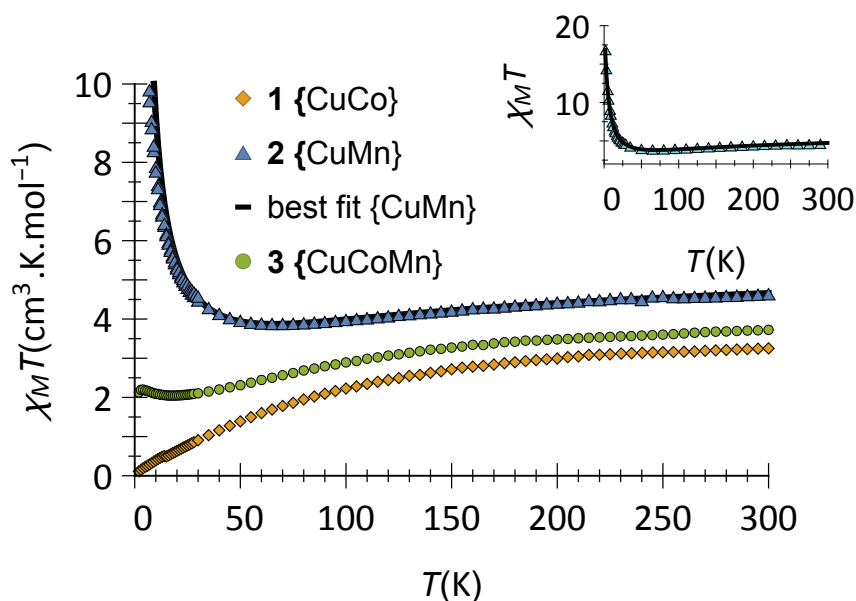


Figure 4. Thermal dependence of the $\chi_M T$ products of 1-3, measured at 5000 Oe ($T > 200$ K) and 200 Oe ($T < 200$ K), the solid line represents the best fit for 2. Inset shows the full temperature range for 2.

In **2**, the presence of a minimum is characteristic of a ferrimagnetic $\{\text{Mn}^{\text{II}}-\text{Cu}^{\text{II}}\}$ chain compound.²⁵⁻²⁷ However, there is no minimum and increase of the $\chi_M T$ product at low temperature for the $\{\text{CuCo}\}$ chain **1**. This is in stark contrast with the oxamato-bridged bimetallic $\{\text{CuCo}\}$ chains reported in the literature.²⁸⁻³¹ For the trimetallic chain **3**, the $\chi_M T$ product steadily decreases but reaches a minimum around 20 K, before going up. We assume that it comes from the presence of Mn^{II} ions with a larger magnetic moment that reinforce the ferrimagnetic character of the chain.

Unfortunately, it is only possible to model the magnetic properties of the $\{\text{CuMn}\}$ chain for which a theoretical alternating chain model has been developed using the following Hamiltonian for the interaction:^{32,33}

$$H = -J \sum_i S_{B_i} \left[(1+\alpha) S_{A_i} + (1-\alpha) S_{A_{i+1}} \right]$$

The least square fitting of the experimental data were done in the 22-300 K temperature range to avoid the effect of interchain interactions. The best fit parameters are equal to $J = -22.0 \text{ cm}^{-1}$, $\alpha = 0.59$ and $g_{\text{Cu}} = 2.14$, with an agreement factor of $4.7 \cdot 10^{-5}$. This corresponds to $J_{\text{oxa}} = -35 \text{ cm}^{-1}$ and $J_{\text{hydro}} = -9.0 \text{ cm}^{-1}$. The largest value is attributed to the interaction between the Cu^{II} and Mn^{II} ions through the oxamato bridge and in accordance with the reported values.^{26,27,34}

There is no model for a $\{\text{CuCo}\}$ alternating chain. However, Drillon *et al.* have developed a model for a regular chain that takes into account spin-orbit coupling.²⁸ In this model –even for small values of the $\text{Cu}^{\text{II}}-\text{Co}^{\text{II}}$ interaction– a minimum of $\chi_M T$ is expected, which gives no satisfactory explanation to the absence of a minimum in **1**. The only possible explanation would be substantial antiferromagnetic interchain interactions, although it seems unlikely owing to the crystal structure

with well isolated chains (shortest interchain metal-metal distances were found at 7.12 Å (Cu-Co) and 7.15 Å (Cu-Mn)).

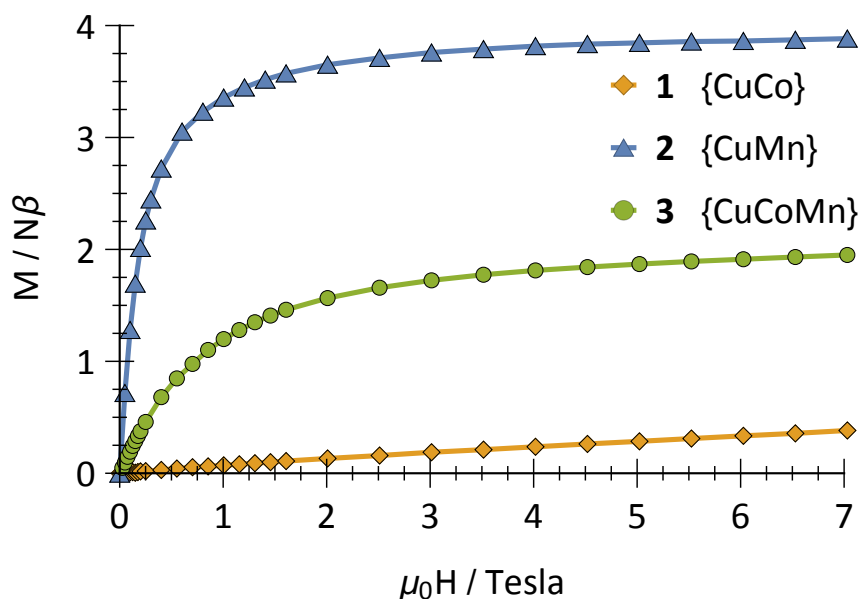


Figure 5. Magnetisation v. field measured at 2 K and between 0 and 7 T for compounds 1-3. The lines are eye guides.

The M versus H plots for compounds 1-3 at 2.0 K are shown in Figure 4. For compound 2 the saturated magnetisation value at 7 T is 3.9 $N\mu_B$, which is consistent for an $S = 2$ ground state for the basic {CuMn} unit resulting from antiferromagnetic coupling between the Mn^{II} and Cu^{II} ions in 2. The M versus H curve of 1 is more puzzling. Even at 7 T, the M value is only 0.38 $N\mu_B$ and the variation of M with the magnetic field is almost linear. It resembles the variation of an antiferromagnetic phase and the presence of a magnetic order would actually be consistent with the evolution of the $\chi_M T$ product with the temperature, which tends towards zero at low temperature. Again, this is puzzling considering the structure of the compound with isolated chains in the crystal. Moreover, since this compound is isostructural with 2, it would be difficult to understand why compound 2 behaves as an isolated ferrimagnetic chain while 1 presents an antiferromagnetic order. To check whether the chains are actually isolated in compound 2, the fit of $\chi_M T$ versus temperature for 2 was repeated on the whole range of temperature including a Curie-Weiss law in the model to take into account the interchain interaction. The best-fit parameters are equal to $J = -23 \text{ cm}^{-1}$, $\alpha = 0.37$, $g_{Cu} = 2.10$ and $\theta = -5.9 \text{ K}$, with an agreement factor of 1.56×10^{-3} . The high negative value for θ in compound 2 supports the hypothesis of a non-negligible antiferromagnetic interchain interaction. The lack of obvious exchange interaction pathways between the chains indicates that the interactions are probably dipolar. However, these non-negligible interactions do not result in a magnetic ordering in the case of compound 2. One plausible explanation is that the interactions develop only in two dimensions. This is consistent with the structure where the chains are organised in layers in the bc plane (see Figure 2). This excludes a magnetic order for compound 2 that contains isotropic Mn^{II} ions while the presence of Ising type Co^{II} ions in 1 could lead to an antiferromagnetic order (see supplementary information)³⁵. Furthermore, this is coherent with the observed magnetisation curves for 3. The introduction of an isotropic ion suppresses the crossover toward an Ising-type behaviour for the

chain, and it prevents the magnetic ordering. The last point to address is the lack of a maximum in the χ_M v. T plot of **1** –i.e. a signature of antiferromagnetic ordering (see supplementary information). The presence of the cobalt hydroxide impurity most likely masks the occurrence of a maximum of χ_M .²⁴

Experimental

All reagents were used as purchased with no further purification. The *N*-(2-hydroxyphenyl)oxamic acid hydrate, mixed sodium salt ($\text{H}_{2.5}\text{Na}_{0.5}\text{-ohpma}\cdot 0.5\text{H}_2\text{O}$) was prepared following the reported procedure.¹⁸

Synthesis.

[Cu(ohpma)Co(OAc)(DMF)₂] (1). $\text{H}_{2.5}\text{Na}_{0.5}\text{-ohpma}\cdot 0.5\text{H}_2\text{O}$ (0.5 mmol, 0.100 g), $\text{Cu(OAc)}_2\cdot\text{H}_2\text{O}$ (0.5 mmol, 0.099 g) and $\text{Co(OAc)}_2\cdot 4\text{H}_2\text{O}$ (0.5 mmol, 0.124 g) were placed in a 23 mL Teflon-lined autoclave reactor filled with DMF (3 mL) and H_2O (1 mL). The reactor was heated to 80°C in 2 h, kept at this temperature for 72 h and then cooled down to room temperature in 24 h. Dark green block crystals of **1** were collected and washed with DMF. Yield: 0.190 g, 89.6% (based on Cu; $M=505.8\text{ gmol}^{-1}$). Elemental analysis (%): calc. for $\text{C}_{16}\text{H}_{21}\text{CuCoN}_3\text{O}_8$: C, 37.99; H, 4.18; N, 8.31; Co, 11.65; Cu, 12.56 (wt. Cu/Co: 1.08). Found: C, 37.83; H, 4.20; N, 8.16; Co, 10.09; Cu, 11.31 (wt. Cu/Co: 1.12). Selected IR data (cm^{-1}): 2932(m), 1627(s), 1566(s), 1473(s), 1370(m), 1344(s), 1297(m), 1274(s), 1242(s), 1110(m), 1027(m), 972(w), 932(w), 886(s), 786(w), 750(s), 675(s), 565(m), 534(m), 373 (m), 325(s).

[Cu(ohpma)Mn(OAc)(DMF)₂] (2). $\text{H}_{2.5}\text{Na}_{0.5}\text{-ohpma}\cdot 0.5\text{H}_2\text{O}$ (0.5 mmol, 0.100 g), $\text{Cu(OAc)}_2\cdot\text{H}_2\text{O}$ (0.5 mmol, 0.099 g) and $\text{Mn(OAc)}_2\cdot 4\text{H}_2\text{O}$ (0.5 mmol, 0.116 g) were placed in a 23 mL Teflon-lined autoclave reactor filled with DMF (4 mL). The reactor was heated to 80°C in 2 h, kept at this temperature for 72 h and then cooled down to room temperature in 24 h. Dark green block crystals of **2** were collected and washed with DMF. Yield: 0.131 g, 62.4% (based on Cu, $M=501.8\text{ gmol}^{-1}$). Elemental analysis (%): calc. for $\text{C}_{16}\text{H}_{21}\text{CuMnN}_3\text{O}_8$: C, 38.29; H, 4.22; N, 8.37; Mn, 10.94; Cu, 12.66 (wt. Cu/Mn=1.16). Found: C, 38.11; H, 4.24; N, 8.19; Mn, 11.72; Cu, 13.53 (wt. Cu/Mn=1.15). Selected IR data (cm^{-1}): 2932(m), 1627(s), 1562(s), 1472(m), 1416(m), 1371(s), 1342(w), 1296(m), 1274(m), 1242(m), 1150(m), 1105(m), 1061(m), 1027(m), 931(w), 838(s), 784(m), 750(m), 670(m), 593 (m), 563(s), 533(m), 446(m), 362(m), 327(m), 285(m).

[Cu(ohpma)Co_{0.57}Mn_{0.43}(OAc)(DMF)₂] (3). $\text{H}_{2.5}\text{Na}_{0.5}\text{-ohpma}\cdot 0.5\text{H}_2\text{O}$ (0.5 mmol, 0.100 g), $\text{Cu(OAc)}_2\cdot\text{H}_2\text{O}$ (0.5 mmol, 0.099 g), $\text{Co(OAc)}_2\cdot 4\text{H}_2\text{O}$ (0.25 mmol, 0.062 g) and $\text{Mn(OAc)}_2\cdot 4\text{H}_2\text{O}$ (0.25 mmol, 0.058 g) were placed in a 23 mL Teflon-lined autoclave reactor filled with DMF (3 mL) and H_2O (1 mL). The reactor was heated to 80°C in 2 h, kept at this temperature for 72 h and then cooled down to room temperature in 24 h. The resulting green solution was then left to crystallise at room temperature in a sealed vial. After 1 day, dark crystals of **3** were collected and washed with DMF. Yield: 0.179 g, 85% (based on Cu, $M=504.1\text{ gmol}^{-1}$). Elemental analysis (%): calc. for $\text{C}_{16}\text{H}_{21}\text{CuCo}_{0.57}\text{Mn}_{0.43}\text{N}_3\text{O}_8$: C, 38.12; H, 4.19; N, 8.33; Co, 6.66; Cu, 12.60; Mn, 4.68 (wt. Cu/Co=1.89, Cu/Mn=2.69; Co/Mn=1.42). Found: C, 37.54; H, 4.01; N, 8.19; Mn, 4.19; Co, 5.99; Cu, 11.44 (wt. Cu/Co=1.91, Cu/Mn=2.73; Co/Mn=1.43). Selected IR data (cm^{-1}): 2930(w), 1628(s), 1565(s), 1472(m), 1371(m), 1343(s), 1296(m), 1274(s), 1242(s), 1109(m), 1027(m), 950(w), 885(s), 785(w), 750(s), 673(s), 564(m), 534(m), 457 (m).

Magnetic measurements.

Magnetic measurements in dc mode were performed on the polycrystalline samples of **1-3** restrained within a capsule with a Quantum Design MPMS SQUID. Magnetic susceptibility data were corrected for the diamagnetism of the constituent atoms using the Pascal's constants. The diamagnetism of the sample holder was measured and subtracted from the raw data.

Elemental analysis.

Elemental analysis was performed in the "service de microanalyse" at ICSN (CNRS, Gif/Yvette, France) and via the PARI program at the IPGP (Paris, France).

FT-IT and TGA.

ATR/FT-IR spectra were collected on a Bruker TENSOR 27 equipped with a simple reflexion ATR diamond plate of the Harrick MPV2 series. The thermogravimetric analysis (**TGA**) was performed on a TA Instruments SDTQ600 under air or nitrogen with a heating rate of 5 °C/min.

Crystallography.

Single-crystal X-Ray diffraction. The data collections for **1-3** were carried out on a Bruker Kappa-APEX II CCD diffractometers. Crystals were mounted on a Hampton Research cryoloop using Paratone-N oil and placed in the cold flow produced with an Oxford Cryosystems device. Data collection strategies were generated with the APEX2 suite of programs (BRUKER).³⁶ The refinement of the unit cell parameters and data reduction were carried out with SAINT (BRUKER),³⁶ and absorptions were corrected with SADABS.^{36,37} The structures were solved with SHELXT-14³⁸ and refined with the SHELXL-2014/7 program³⁹ in the OLEX2 software.⁴⁰ Crystallographic parameters and data refinement results for **1-3** are given in Table 2. CCDC-2032124-2032126 contain the supplementary crystallographic data for this paper. The data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/structures

Powder X-ray diffraction. The X-ray powder diffraction (XRD) patterns were collected at the X-ray diffraction platform of the IMPMC, from samples deposited on zero background silicon holders, using an X'Pert Pro MPD Panalytical diffractometer in Bragg-Brentano geometry equipped with a Co-K α radiation source ($\lambda_{K,1} = 1.78897 \text{ \AA}$, $\lambda_{K,2} = 1.79285 \text{ \AA}$) and an X'Celerator detector. Le Bail refinements⁴¹ were performed with the FullProf suite of programs.⁴²

	1	Powder XRD Le Bail refinement		2	Powder XRD Le Bail refinement		3	Powder XRD Le Bail refinement	
Formula	C ₁₆ H ₂₁ CuCoN ₃ O ₈			C ₁₆ H ₂₁ CuMnN ₃ O ₈			C ₁₆ H ₂₁ CuCo _{0.78} Mn _{0.22} N ₃ O ₈		
FW [g mol ⁻¹]	505.83			501.84			504.94		
Crystal system	Monoclinic	Monoclinic		Monoclinic	Monoclinic		Monoclinic	Monoclinic	
Space group	<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> 2 ₁ / <i>c</i>		<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> 2 ₁ / <i>c</i>		<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> 2 ₁ / <i>c</i>	
<i>a</i> [Å]	15.3056(4)	15.5124(3)		15.1715(7)	15.3346(2)		15.2626(4)	15.4925(2)	
<i>b</i> [Å]	8.3802(2)	8.4173(1)		8.4796(4)	8.5205(1)		8.4325(2)	8.44424(8)	
<i>c</i> [Å]	15.8983(4)	15.9123(4)		16.1749(7)	16.1908(2)		16.0027(4)	15.9995(5)	
β [°]	98.204(1)	97.997(2)		98.443(3)	98.372(2)		98.2690(10)	98.085(2)	
<i>V</i> [Å ³]	2018.31(9)	2057.51(7)		2058.32(16)	2091.92(4)		2038.17(9)	2064.78(7)	
<i>Z</i>	4			4			4		
<i>T</i> [K]	200(2)	298		200(2)	298		200(2)	298	
λ [Å]	1.54178 (Cu-K α)	1.789190 (Co-K α)		0.71073 (Mo-K α)	1.789190 (Co-K α)		1.54178 (Cu-K α)	1.789190 (Co-K α)	
ρ_{calc} [g cm ⁻³]	1.665			1.619			1.646		
μ [mm ⁻¹]	8.180			1.694			7.803		
<i>F</i> (000)	1032			1024			1030		
Crystal size [mm]	0.15x0.12x0.08			0.12x0.11x0.08			0.22x0.2x0.12		
<i>q</i> limits [°]	5.841-67.725	4.111-70		1.357-30.095	4.111-70		5.859-68.294	4.111-59.99	
Measured reflections	27128			17547			17152		
Unique reflections	3637			6040			3725		
<i>R</i> _{int}	0.034			0.053			0.0309		
Reflections <i>I</i> >2 σ (<i>I</i>)	3541			3797			3437		
Parameters	267			267			268		
Restraints	0			0			0		
<i>R</i> ₁ [<i>I</i> >2 σ (<i>I</i>)] / <i>R</i> _p	0.0267	0.122		0.0412	0.136		0.0277	0.113	
<i>wR</i> ₂ / <i>R</i> _{wp}	0.0714	0.147		0.0829	0.109		0.0696	0.121	
GOF	1.018			0.969			1.047		
Largest residuals [eÅ ⁻³]	-0.547;0.338			-0.359;0.418			-0.261;0.363		

$$R_1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|} \quad R_p = \frac{\sum |y_o - y_c|}{\sum y_o} \quad \omega R_2 = \sqrt{\frac{\sum (\omega(F_o^2 - F_c^2)^2)}{\sum (\omega(F_o^2)^2)}} \quad R_{wp} = \sqrt{\frac{\sum (\omega(y_o^2 - y_c^2))}{\sum (\omega(y_o^2)^2)}}$$

Table 2. Crystallographic parameters and data refinement for compound 1-3.

Conclusions

New neutral heterometallic oxamate-based 1D compounds have been successfully synthesized from an *ortho*-substituted phenyloxamate ligand with an hydroxido group. They represent the first examples of 1D oxamate-based coordination polymers obtained solvothermally, and more generally, the first examples of compounds with fully coordinating oxamate ligands prepared solvothermally. In comparison to the multi-step approach traditionally used for oxamate chemistry, the one-pot method used here has not been an issue for the preparation of heterobimetallic compounds. It has even been possible to prepare a heterotrimetallic chain. This is made possible by the strong affinity of copper ions for the $\kappa N, \kappa O$ chelating site of the oxamate ligands. As such, copper-based heterometallic systems represent ideal targets for the solvothermal preparation of oxamate-based magnetic coordination polymers.

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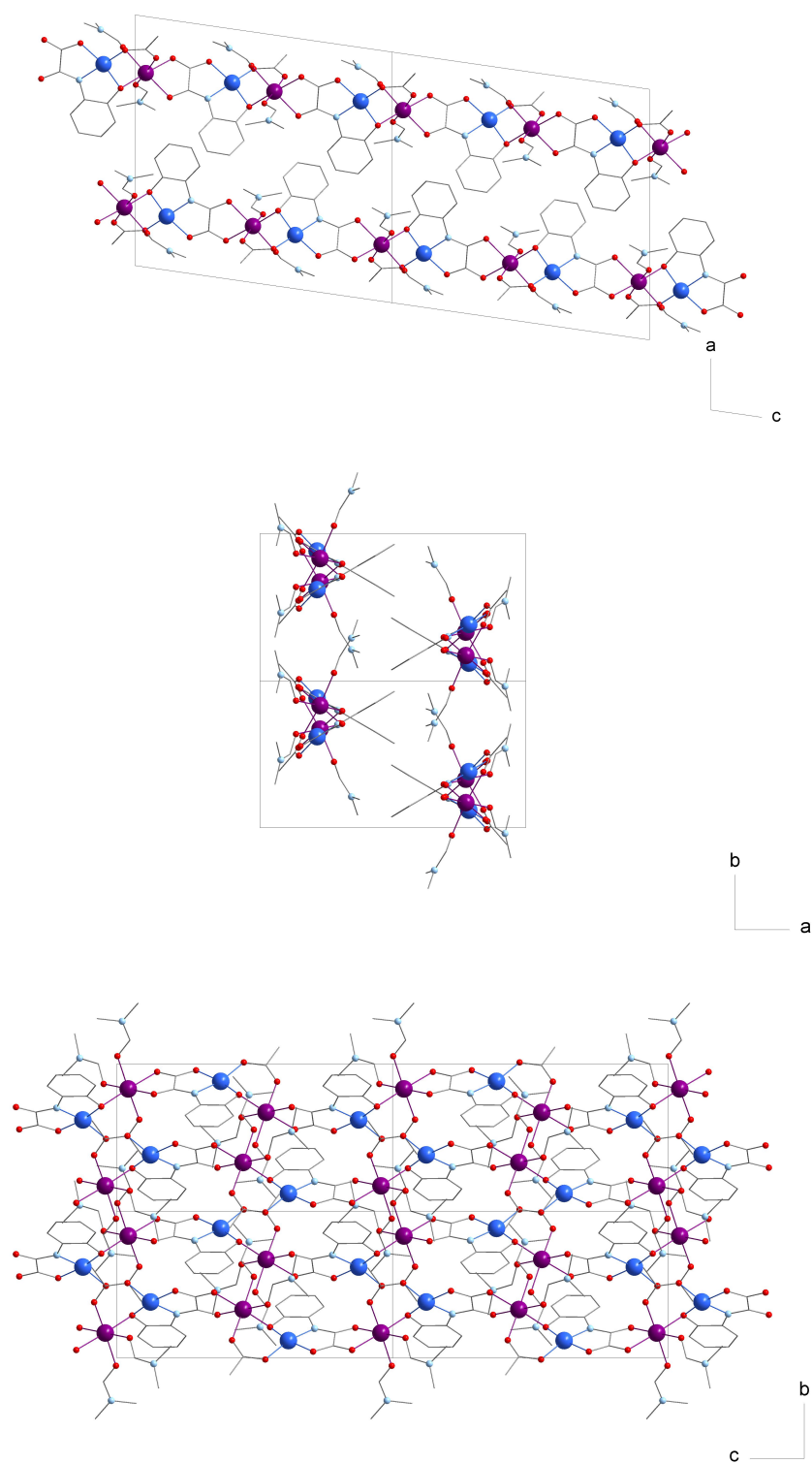
Scholarship Council (CSC) with M. Ang Li's PhD fellowship. Elemental analysis was supported by the IPGP multidisciplinary program PARI, and by Paris–IdF region SESAME Grant no. 12015908.

Notes and references

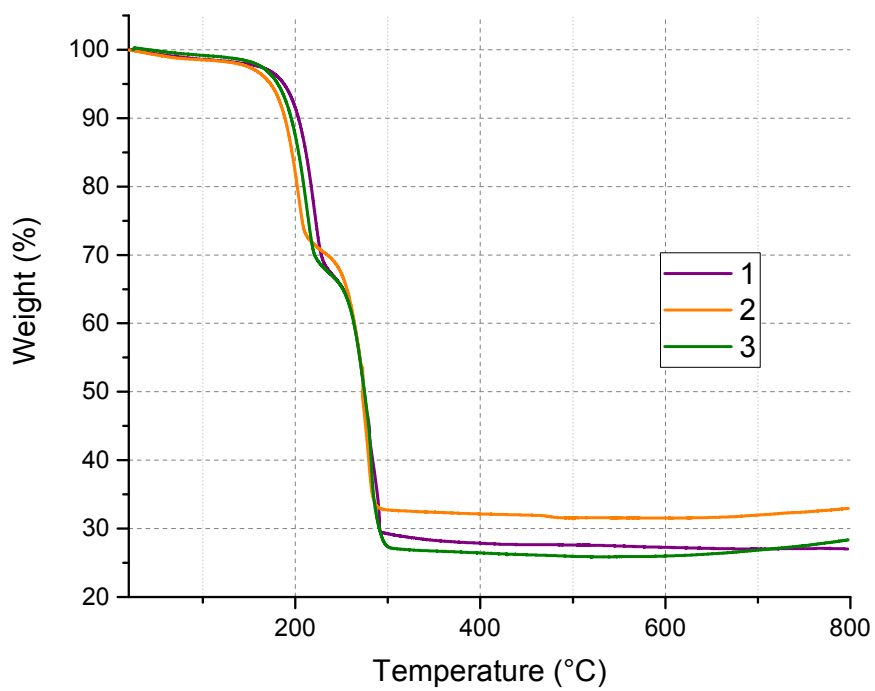
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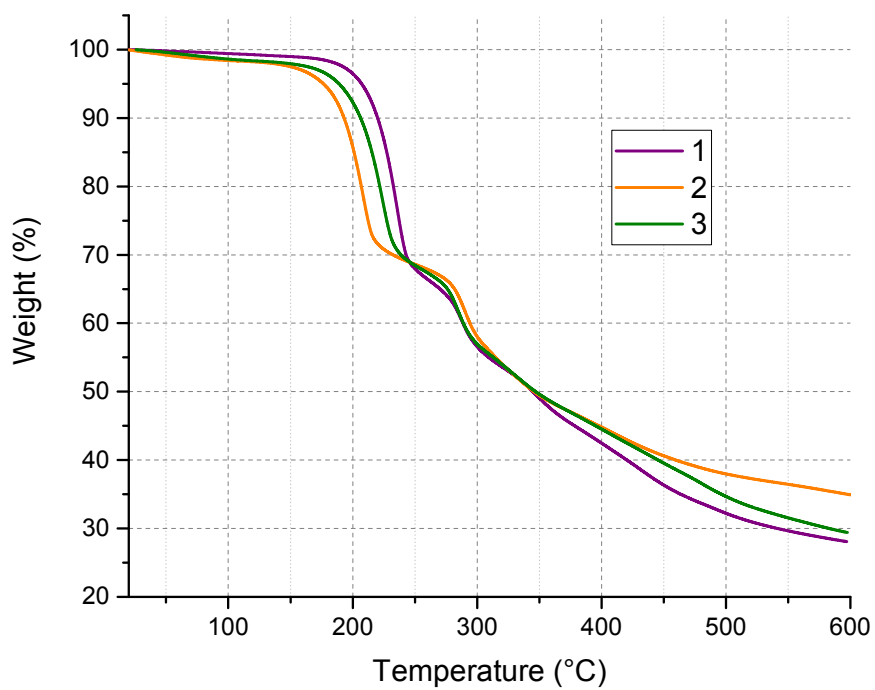
Supplementary information



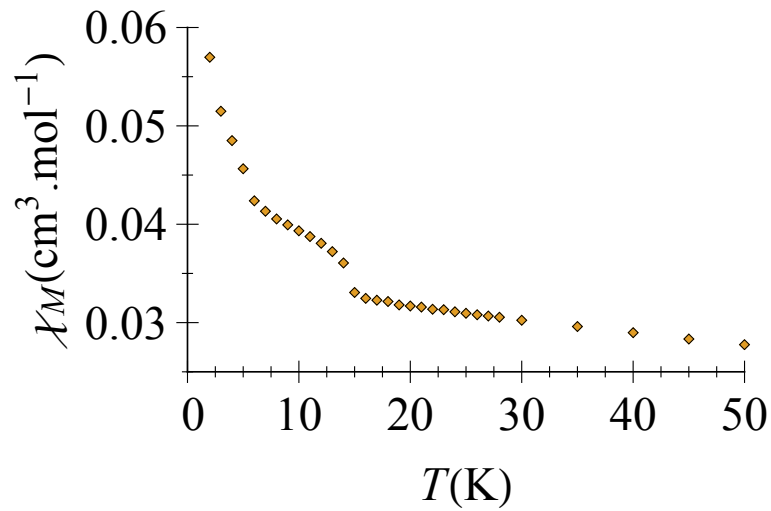
Crystal packing of 1.



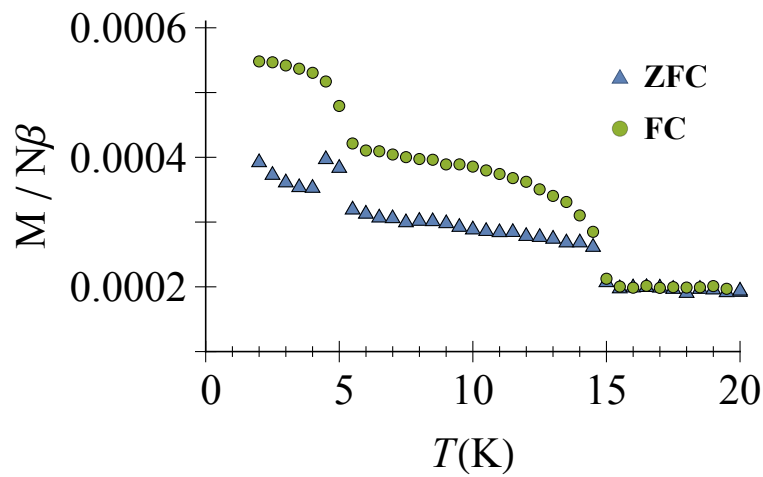
TGA of compounds 1-3 performed under air at 5°C/min.



TGA of compounds 1-3 performed under N₂ at 5°C/min.



Thermal dependence of the χ_M of 1 measured at 200 Oe



ZFC/FC curves measured at 30 Oe for 1.

	Ising	XY	Heisenberg
1D	No	No	No
2D	Yes	KT transition ⁴³	No
3D	Yes	Yes	Yes

Magnetic ordering v. dimensionality of a network and magnetic anisotropy.