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Visible, Near-Infrared, and Dual-Range Luminescence Spanning the 4f Series Sensitized by a Gallium(III)/Lanthanide(III) Metallacrown Structure

Elvin V. Salerno,[‡] Svetlana V. Eliseeva,^{*,‡} Bernadette L. Schneider, Jeff W. Kampf, Stéphane Petoud,^{*} and Vincent L. Pecoraro^{*}

ABSTRACT: Lanthanide(III) ions (Ln^{3+}) in coordination compounds exhibit unique luminescent properties with narrow and characteristic f-f transitions throughout the visible and near-infrared (NIR) ranges. In addition, some Ln^{3+} such as Pr^{3+} , Sm^{3+} , Dy^{3+} , Ho^{3+} , Er^{3+} and Tm^{3+} possess an exceptional ability, although less explored, to exhibit dual-range emissions. Such remarkable features allow highly specific use in materials sciences and biology, for example, for the creation of sophisticated barcode modules or for the next generation of optical imaging applications. Herein, a series of $\text{Ga}^{3+}/\text{Ln}^{3+}$ metallacrowns (MCs) with the general composition $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]\text{Na}\cdot x\text{CH}_3\text{OH}\cdot y\text{H}_2\text{O}$ (**Ln-1**, $\text{Ln} = \text{Pr}^{3+}$, Nd^{3+} , Sm^{3+} – Yb^{3+} and analogue Y^{3+} ; $\text{H}_3\text{shi} = \text{salicylhydroxamic acid}$) is presented. **Ln-1** were obtained by reacting Ga^{3+} and Ln^{3+} nitrate salts with the H_3shi ligand. X-ray single crystal unit cell analysis confirmed that all MCs are isostructural. The crystal structure was solved for the Nd^{3+} analogue and revealed that Nd^{3+} is centered between two $[\text{12-MC}_{\text{Ga}^{\text{III}}\text{N}(\text{shi})-4}]$ MC rings and bound to eight hydroximate oxygen ions (four from each ring) in a pseudo-square antiprismatic fashion adopting a pseudo- D_{4h} symmetry. PGSE DOSY ^1H NMR spectroscopy and ESI mass-spectrometry confirmed that the structure of **Ln-1** remains intact in methanol solutions while mass spectrometry suggests that four OH^- bridges are exchanged with $\text{CH}_3\text{O}^-/\text{CD}_3\text{O}^-$. An exceptional ability of this series of MCs to sensitize the characteristic emission of Ln^{3+} was confirmed with the observation of bright red and green emission signals of **Eu-1** and **Tb-1**, NIR emissions of **Yb-1** and **Nd-1**, and dual-range emissions of **Pr-1**, **Sm-1**, **Dy-1**, **Ho-1**, **Er-1** and **Tm-1** in the solid state upon excitation into ligand-centered bands at 340 nm. The luminescence properties of **Ln-1** ($\text{Ln} = \text{Nd}^{3+}$, Sm^{3+} , Eu^{3+} , Tb^{3+} , Dy^{3+} , Yb^{3+}) were also investigated in CH_3OH and CD_3OD solutions. For **Eu-1** and **Yb-1** MCs, more extensive analyses of the photophysical properties were performed that included the determination of radiative lifetimes, intrinsic quantum yields and sensitization efficiencies. The absolute quantum yields (Q_{Ln}^{L}) of **Ln-1** in the visible and in the NIR ranges have been determined. In the case of **Sm-1** the values of Q_{Ln}^{L} in CH_3OH and CD_3OD solutions are exceptionally high, i.e. 10.1(5) % and 83(1) %. Values obtained for **Yb-1**, i.e. 0.78(4) % in CH_3OH and 8.4(1)% in CD_3OD , are among the highest ones reported today for Yb^{3+} complexes formed with non-deuterated and non-halogenated ligands.

INTRODUCTION

Metallacrowns (MCs) constitute a class of supramolecular complexes which possess a $[\text{M-N-O}]_n$ repeating ring motif.¹⁻² The M^{n+} ring cations are usually 3d transition metal ions (from vanadium(V) to zinc(II)), but MCs with the group 13 triels such as gallium(III)³⁻⁶ or with 4d and 5d transition metal ions (e.g. palladium(II), platinum(II))⁷ have also been synthesized. Due to the diversity of structures and functional properties combined with the ease of synthesis via self-assembly in (typically) ambient temperature conditions MCs have found a broad range of applications in different fields such as molecular recognition,⁸⁻¹⁴ molecular magnetism,¹⁵⁻²³ and magnetorefrigeration,²⁴ as well as contrast agents for magnetic resonance imaging (MRI)²⁵⁻²⁷ or as luminescent probes for optical imaging.²⁸⁻²⁹

It has been shown that many MCs are able to bind a single or multiple trivalent lanthanide(III) ions (Ln^{3+}).¹ Ln^{3+} are widely studied for their unique optical properties derived from core-like valence 4f orbitals. Since 4f orbitals are present within the 5s and 5p orbitals, they are largely shielded from crystal field and covalent bonding effects. Such electronic structure results in Ln^{3+} displaying narrow

and characteristic in their absorption and emission spectra from the corresponding f-f transitions, the positions of which are generally not significantly affected by the coordination environment or environmental conditions (e.g. pH, temperature, biomolecules).³⁰ The characteristic emission of Ln^{3+} compounds that cover UV, visible, and near-infrared (NIR) ranges lead to their various applications in materials science³¹ and biology:³²⁻³³ from telecommunications,³⁴ secret tags,³⁵ lighting, display technologies,³⁶ and night-vision devices³⁷ to optical imaging and sensing.³⁸⁻⁴¹ Although Ln^{3+} luminescence is widely used, f-f transitions suffer from low molar absorptivities ($<10 \text{ M}^{-1}\text{cm}^{-1}$) due to the Laporte- and spin- forbidden nature of most of such transitions. This forbiddance leads to experimentally useful long-lived excited state lifetimes of Ln^{3+} , but, it implies that Ln^{3+} emission must be sensitized indirectly in order to obtain a significant emission intensity.³⁰ One of the strategies for Ln^{3+} sensitization is based on the ‘antenna effect’⁴² that is typically achieved by surrounding Ln^{3+} with organic chromophoric ligands. As a general scheme, the energy is absorbed by the ligand through singlet states ($\text{S}_0 \rightarrow \text{S}_n$), followed by internal conversion ($\text{S}_n \rightarrow \text{S}_1$) and intersystem crossing from excited singlet to excited triplet states

($S_1 \rightarrow T_1$), then, in most of the cases, transfer from these excited states to the accepting electronic levels of Ln^{3+} ($S_1 \rightarrow \text{Ln}^{3+*}$, $T_1 \rightarrow \text{Ln}^{3+*}$).^{30, 43-44} Other processes such as metal to ligand charge transfer (MLCT) or inter- and intraligand charge transfer (ILCT) states may also play a role in the Ln^{3+} sensitization process.⁴⁵⁻⁵⁰ Another challenge that has to be addressed when creating luminescent Ln^{3+} -based coordination compounds is the protection of the Ln^{3+} against sources of nonradiative deactivations which are most commonly due to overtones of O-H, N-H, and C-H vibrations.⁵¹ The probability of this process is inversely proportional to the energy gap (ΔE) between the emitting level of the Ln^{3+} and the highest energy level of the ground multiplet. Thus, Eu^{3+} (12 300 cm^{-1}) and Tb^{3+} (14 800 cm^{-1}), having the largest ΔE (not considering UV-emitting Gd^{3+}), exhibit bright red and green emissions in their coordination compounds, and are the most studied and exploited.⁵²⁻⁵⁴ In respect to the NIR emission range, Yb^{3+} and Nd^{3+} complexes have been mainly considered.³⁸ Other luminescent Ln^{3+} , i.e. Pr^{3+} , Sm^{3+} , Dy^{3+} , Ho^{3+} , Er^{3+} and Tm^{3+} , usually exhibit less intense emission signals partially because of a smaller ΔE or due to the complex electronic structures with several emitting levels.⁵⁵ However, these Ln^{3+} possess a very unique ability to emit in both the visible and in the NIR ranges.^{47, 50, 56} Such dual-emission features of Ln^{3+} are less studied and exploited for coordination compounds⁵⁷⁻⁶² but can lead to, for example, the creation of sophisticated barcode modules⁶³⁻⁶⁶, or novel optical imaging applications.⁶⁷⁻⁶⁸

We have recently reported several series of visible and NIR emitting $\text{Ln}^{3+}/\text{Ga}^{3+}$ MCs^{3, 5-6, 69} and metallocryptates⁷⁰ assembled using salicylhydroxamic acid (H_3shi) as building blocks. We had demonstrated that $\text{Ga}^{3+}/\text{shi}$ scaffolds possess a remarkable ability to sensitize the characteristic luminescence of bound visible and NIR-emitting Ln^{3+} including the dual-emitting Sm^{3+} and Dy^{3+} . However, quantitative photophysical properties (i.e. total ($Q_{\text{Ln}}^{\text{Ln}}$) and intrinsic ($Q_{\text{Ln}}^{\text{Ln}}$) quantum yields, observed (τ_{obs}) and radiative (τ_{rad}) lifetimes, sensitization efficiencies (η_{sens}) have, despite the similarities between the $\text{Ga}^{3+}/\text{shi}$ based scaffolds, been largely and in many cases unpredictably affected by the nature of the bridging ligands, counter-cations, crystals packing etc.

Herein we present the synthesis, characterization (including X-ray single crystal structural analysis) and detailed photophysical studies of a series of $\text{Ga}^{3+}/\text{Ln}^{3+}$ MCs with the general composition $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]\text{Na}\cdot x\text{CH}_3\text{OH}\cdot y\text{H}_2\text{O}$ (**Ln-1**, $\text{Ln} = \text{Pr}^{3+}-\text{Yb}^{3+}$, excluding Pm^{3+}). The Y^{3+} analogue was also synthesized to facilitate characterization by ^1H NMR spectroscopy. Steady-state excitation and emission spectra collected in the visible and/or the NIR ranges, $Q_{\text{Ln}}^{\text{Ln}}$ and τ_{obs} were measured and analyzed for **Ln-1** in the solid state and, when possible, for solutions in CH_3OH and CD_3OD . The phosphorescence spectrum of **Gd-1** in the solid state was acquired to determine the energy position of the ligand triplet state. More detailed analyses of the photophysical properties, including the determination of τ_{rad} , $Q_{\text{Ln}}^{\text{Ln}}$ and η_{sens} , were performed for **Eu-1** and **Yb-1**.

The presented series of **Ln-1** MCs is remarkable relative to the $\text{Ga}^{3+}/\text{Ln}^{3+}$ MCs that we have published previously.^{3, 5-6, 69} In particular, Ln^{3+} in **Ln-1** adopt a pseudo- D_{4h} symmetry compared to C_4 or C_1 symmetries of our previously described MC systems. In addition, **Ln-1** possess a unique ability to sensitize the visible, NIR or dual-emissions of ten Ln^{3+} of different natures. The present structure is also the first MC in which efficient sensitization of red Eu^{3+} emission occurs. We should note that a $\text{Ga}^{3+}/\text{Dy}^{3+}$ MC possessing a molecular geometry which shares structural similarities with **Ln-1** but which is formed with another counter-cation ($n\text{Bu}_4\text{N}^+$ vs. Na^+) was obtained through a different synthetic route and was independently reported by the Rentschler group in a recent communication focused exclusively on single molecule magnet properties.⁷¹ Luminescent properties were not described in that publication. It is important to note that we have previously evidenced that the nature of the counter-cation in the MC structure may significantly affect photophysical properties.⁶⁹

METHODS

Excitation, emission spectra, observed lifetimes and quantum yields. Luminescence data were collected for **Ln-1** samples in the solid state or on freshly prepared 50 μM solutions in CH_3OH and CD_3OD placed into 2.4 mm i.d. quartz capillaries or quartz Suprasil cells. Emission and excitation spectra were measured on a Horiba-Jobin-Yvon Fluorolog 3 spectrofluorimeter equipped with either a visible photomultiplier tube (PMT) (220–800 nm, R928P; Hamamatsu), a NIR solid-state InGaAs detector cooled to 77 K (800–1600 nm, DSS-IGA020L; ElectroOptical Systems, Inc., USA), or a NIR PMT (950–1650 nm, H10330-75; Hamamatsu). All spectra were corrected for the instrumental functions. Luminescence lifetimes were determined under excitation at 355 nm provided by a Nd:YAG laser (YG 980; Quantel), the signals in the visible and the NIR ranges were detected with a R928 or H10330-75 PMTs connected to an iHR320 monochromator (Horiba Scientific). The output signals from the detectors were fed into a 500 MHz bandpass digital oscilloscope (TDS 754C; Tektronix) and transferred to a PC for data processing with the Origin 8@ software. Luminescence lifetimes are averages of three or more independent measurements. Quantum yields were determined with a Fluorolog 3 spectrofluorimeter based on the absolute method using an integration sphere (GMP SA, Renens, Switzerland). Each sample was measured several times. The experimental error for the determination of quantum yields is estimated as $\sim 10\%$.

Phosphorescence spectrum. Phosphorescence spectrum of **Gd-1** in the solid state at 77 K was acquired on a Horiba-Jobin-Yvon Fluorolog 3 spectrofluorimeter in time-resolved mode. See Supporting Information for more details.

Diffuse reflectance spectra. For collection of diffuse reflectance spectra **Ln-1** (5 wt.%) were thoroughly grinded with MgO . Measurements were performed on a Jasco V670 UV-visible spectrophotometer using a horizontal integration sphere accessory at room temperature.

Absorption spectra. Absorption spectra (Figure S9) were collected on freshly prepared 50 μM solutions of **Ln-1** in methanol (Merck, Uvasol®) placed in quartz cuvettes on a Jasco V670 UV-visible spectrophotometer in absorbance mode.

Synthesis and characterization. All reagents and chemicals were purchased from commercial sources and used without further purification. All reactions were carried out aerobically under ambient conditions. Elemental analysis was performed by Atlantic Microlabs Inc. ESI-MS spectra were collected with an Agilent 6230 TOF HPLC-MS mass spectrometer in negative ion mode (-350V) on samples dissolved in methanol or acetonitrile at a concentration of 2 mg/mL. Single crystals suitable for x-ray analysis were grown via slow evaporation from methanol/water solution at 295 K. See supporting information for x-ray analysis details.

Salicylhydroxamic acid (H_3shi , 0.4 mmol), $\text{Ln}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$ (0.05 mmol, $\text{Ln}^{3+} = \text{Pr}^{3+}, \text{Nd}^{3+}, \text{Sm}^{3+}, \text{Eu}^{3+}, \text{Gd}^{3+}, \text{Tb}^{3+}, \text{Dy}^{3+}, \text{Ho}^{3+}, \text{Er}^{3+}, \text{Tm}^{3+}, \text{Yb}^{3+}, \text{Y}^{3+}$), and $\text{Ga}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$ (0.4 mmol) were dissolved in 40 mL methanol. Concentrated (~19.8 M) aqueous NaOH solution (1.4 mmol) was added and the solution was stirred overnight. The solution was filtered and the filtrate was left for slow evaporation, producing crystalline compound within 2-3 weeks. Compound was collected via filtration and dried in air. See supporting information for details.

NMR. One and two dimensional ^1H NMR were performed on **Y-1** using a Varian VNMRs 500 MHz spectrometer equipped with a Varian 5 mm PFG OneNMR Probe. See supporting information for details.

RESULTS

Synthesis and crystal structure of Ln-1

The reaction between $\text{Ga}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$, $\text{Ln}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$ ($\text{Ln} = \text{Pr}^{3+}, \text{Nd}^{3+}, \text{Sm}^{3+}-\text{Yb}^{3+}$) and salicylhydroxamic acid (H_3shi) in a concentrated aqueous sodium hydroxide solution led to the formation of **Ln-1** with the general formula $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]_n\text{Na} \cdot x\text{CH}_3\text{OH} \cdot y\text{H}_2\text{O}$. Solids were collected as crystalline materials and characterized by ESI mass spectrometry, elemental analysis, and single crystal unit cell analysis.

X-ray quality crystals of **Nd-1** were grown from the slow evaporation of a methanol solution at ambient temperature. All compounds were found to have identical unit cells confirming the isostructural nature of the compounds (Table S2) allowing a global description of molecular structures of this family of MCs.

The crystal structure of **Nd-1** is depicted in Figure 1d. It was solved in a monoclinic $P2_1/c$ group. Each molecule contains two $[12\text{-MC}_{\text{Ga}^{\text{III}}\text{N}(\text{shi})\text{-4}}]$ units with composition that are identical to the MCs presented previously (Figure 1b, c).³⁻⁵ Each of these MC units is slightly concave in nature and consists of four shi^{3-} ligands bridging four Ga^{3+} ions to form a neutral ring. The Nd^{3+} ion is located in the center of the two MC rings being bound to eight hydroximate oxygen ions (four from each ring) in a pseudo-square

antiprismatic fashion. Each molecular unit contains a pseudo D_{4h} symmetry about the Nd^{3+} . Each of the four MC Ga^{3+} ions are connected to a counterpart MC Ga^{3+} ion via a bidentate hydroxide $[\text{Ga}(\mu_2\text{OH})\text{-Ga}]$ bridge. The overall charge of the **Ln-1** MC is -1 , balanced by a Na^+ cation.

There are two distinct complex structures present in the crystal of **Nd-1**. The first contains two *C* (clockwise) MC rings of $[-\text{Ga-N-O-}]_4$ motif (when viewed with Nd^{3+} ion on the opposite side of the ring) binding a single Nd^{3+} , while the second complex contains two *A* (anti-clockwise) MC rings of $[-\text{Ga-O-N-}]_4$ motif binding a single Nd^{3+} (Figure S3). The nature of these complexes is consistent with an enantiomeric (non-superimposable mirror-image) relationship between the two isomers such that the above structural description is applicable to both of them. Their physical properties will generally be identical.

Charge neutrality is maintained, and the two isomers are bridged by two Na^+ such that discrete molecular dimers are present in the crystal (Figure S4). Each Na^+ is bound to a phenoxo oxygen and a hydroximate oxygen belonging to each of the two isomers. Each point of contact is additionally bonded to a Ga^{3+} ion in a $[\text{Ga}(\mu_2\text{O})\text{-Na}]$ organization. Two oxygens from solvent molecules complete the distorted six-coordinate geometry about the Na^+ ion. Different types of hydrogen bonds formed with solvents of crystallization (water and methanol) can be observed in the structure.

Behavior of Ln-1 in solution

To assess the behavior of the complexes in solution, mass spectrometry experiments were performed within a timeframe of 10–30 min after preparation of solutions of **Ln-1** in methanol and in acetonitrile. For each Ln^{3+} analogue in acetonitrile solution, a single large peak with isotopic distribution was identified (except for **Nd-1** and **Pr-1**, for which an additional feature was detected (ESI+)), which was consistent with the molecular formula of the monoanion $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]^-$ ($[\text{M}]^-$). This result suggests that the crystalline materials of **Ln-1** are isolated as $\text{Na}[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]$.

For most Ln^{3+} analogues analyzed by mass spectrometry in methanol solution, the base species $[\text{M}]^-$ were also identified, however, spectra were consistent with the presence of four additional species, i.e. $[\text{M}+14]^-$, $[\text{M}+28]^-$, $[\text{M}+42]^-$, and $[\text{M}+56]^-$. These can be assigned to a sequential replacement of the bridging hydroxide ions (OH^-) with methoxide ions (CH_3O^-) as in the formula: $[\text{LnGa}_8(\text{shi})_8(\text{OH})_{4-n}(\text{CH}_3\text{O})_n]^-$. This observation is in agreement with the exchange of CH_3O^- for OH^- from the gallium(III) ring metals across the entire series of Ln^{3+} MCs. The short timeline between the solution preparation (maximum 30 min) and the mass spectrometry experiment suggests that $\text{CH}_3\text{O}^-/\text{OH}^-$ exchange occurs rapidly in methanol solution. Over time $[\text{LnGa}_8(\text{shi})_8(\text{CH}_3\text{O})_4]^-$ should become the dominant anion in solution given the prevalence of available CH_3O^- vs. OH^- (in a 2 mM solution of **Ln-1** in methanol, the methanol concentration is 24.7 M) although differences in pKa between HO^- and CH_3O^- will also be relevant.⁷²

To examine the solution stability further, we performed Pulsed Gradient Spin Echo Diffusion Ordered ^1H NMR Spectroscopy (PGSE DOSY) on a diamagnetic Ln^{3+} analogue (**Y-1**) in CD_3OD .⁷³ NMR signal response to increasing field gradient strength was fit quantitatively to determine experimental diffusion coefficients based on previously described procedures (Table S3).⁷⁴⁻⁷⁵ In this experiment, it was found that all the identified MC protons migrate together within the time resolution of our experiment. In particular, no protons were found to migrate with a different diffusion coefficient which would be an indication of decomposition, of the MC. Calculations of the diffusion coefficient (D) based on the molecular weight and the Stokes-Einstein equation further support the stability of

the MC dimer in solution (Figures S7, S8).⁷³ For the **Y-1** complex, a D value of $4.10 \cdot 10^{-10} \text{ m}^2\text{s}^{-1}$ was calculated. For a hypothetical monomer of **Y-1**, we calculated $D = 4.54 \cdot 10^{-10} \text{ m}^2\text{s}^{-1}$. Finally, for a MC dimer with similar ethynylsalicylhydroximate (eshi³⁻) ring ligands but a larger bridging ligand, i.e. isophthalate (iph⁻), and two Ln^{3+} , ($[\text{Ln}_2\text{Ga}_8(\text{eshi})_8(\text{iph})_4]$), we found $D = 3.17 \cdot 10^{-10} \text{ m}^2\text{s}^{-1}$.⁷⁶ The experimental D value for **Y-1**, $4.23(8) \cdot 10^{-10} \text{ m}^2\text{s}^{-1}$, is mostly similar to the calculated value for $\text{NaLnGa}_8(\text{shi})_8(\text{OH})_4$, and is an intermediate between the values calculated for the monomer and the larger dimer. This suggests that **Y-1** exists in solution as the $[\text{YGa}_8(\text{shi})_8(\text{OH})_{4-n}(\text{CH}_3\text{O})_n]$ complex in agreement with the crystal structure and mass spectrometry results.

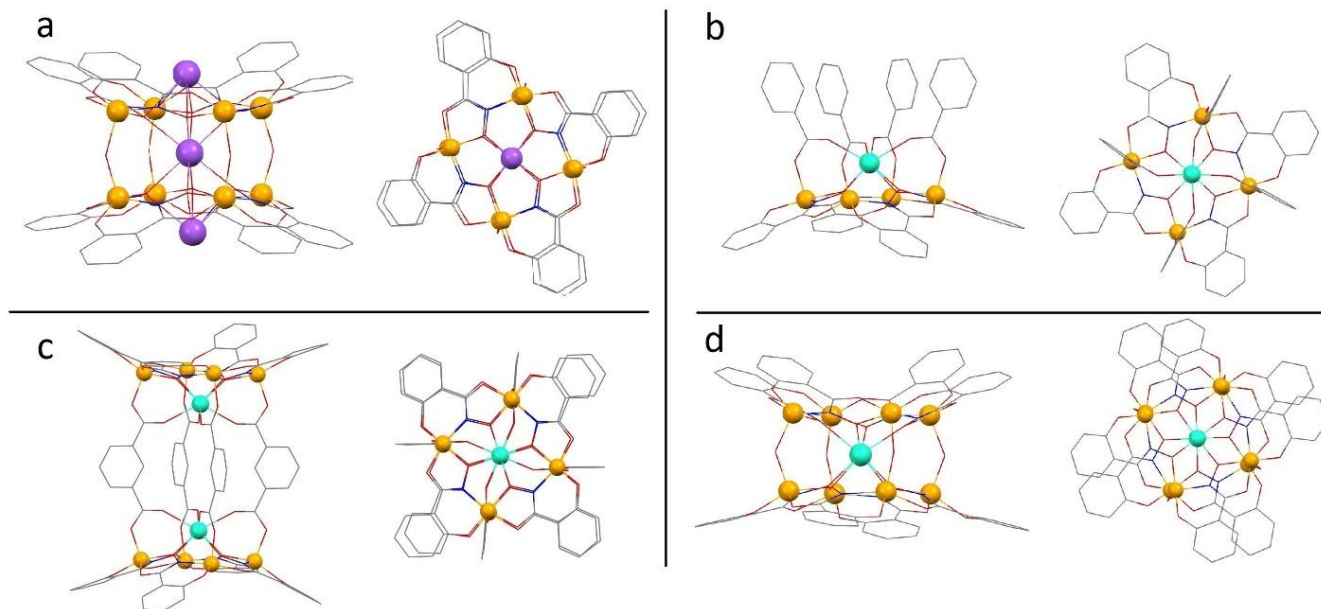


Figure 1. Side (left) and top-down (right) views of $\text{Ga}^{3+}/\text{shi}$ MCs developed by our group. (a) $[\text{Na}_3\text{Ga}_8(\text{shi})_8(\text{OH})_4]_4$ (b) $[\text{DyGa}_4(\text{shi})_4(\text{benzoate})_4]_5$ (c) $[\text{Dy}_2\text{Ga}_8(\text{shi})_8(\text{isophthalate})_4]_3$ (d) $[\text{NdGa}_8(\text{shi})_8(\text{OH})_4]$. Solvents of crystallization, non-integral counter ions, and hydrogen atoms have been omitted for clarity. Color code: Ga, orange; Na, violet; Ln, teal; O, red; N, blue; C, grey.

Photophysical properties

Ligand-centered photophysical properties

Absorption spectra of **Ln-1** collected in methanol solution (Figure S9) are similar for all the MCs formed with the different lanthanides of and present broad bands in the UV region extending up to 350 nm that are attributed to $\pi \rightarrow \pi^*$ transitions. Apparent maxima of the lowest in energy absorption bands are located at 313–314 nm and have molar absorption coefficients of $\sim 3.6 \cdot 10^4 \text{ M}^{-1} \text{ cm}^{-1}$. The energy of the ligand singlet state (S_1) determined from the edge of the absorption spectrum is located at $28\,690 \text{ cm}^{-1}$ (348.5 nm). Diffuse reflectance spectra of **Ln-1**, similarly to absorption spectra, are dominated by broad-band $\pi \rightarrow \pi^*$ transitions in the UV range up to 370 nm (Figure S10). Sharp features attributed to f-f transitions have very weak contributions if diffuse reflectance spectra are presented as Kubelka-Munk function vs. wavelength. All of the **Ln-1** complexes exhibit similar diffuse reflectance spectra ex-

cept for **Eu-1** which has a small but distinct extension to the red side of the lowest energy band at $\sim 375 \text{ nm}$ (Figure S10, right). This difference can likely be attributed to the lower reduction potential of $\text{Eu}^{3+}/\text{Eu}^{2+}$ in respect to other Ln^{3+} and to the formation of a ligand-to-metal charge transfer (LMCT) state.

Gd^{3+} complexes serve as useful probes to assess the electronic structure of a ligand bound to Ln^{3+} due to the Gd^{3+} possessing an accepting energy level ($32\,000 \text{ cm}^{-1}$)⁷⁷ that is too high to be populated by commonly used chromophoric organic ligands. Such complexes are therefore useful to assign ligand triplet (T_1) states energies as this state is not depopulated by Gd^{3+} . Moreover, its heavy-atom and paramagnetic effects facilitate intersystem crossing⁷⁸ and make phosphorescence emission arising from this T_1 state more probable and, in turn, more intense. The energy position of the T_1 level ($22\,620 \text{ cm}^{-1}$ or 442 nm) was determined as the 0-0 transition from the phospho-

rescence spectrum of the **Gd-1** in the solid state recorded in time-resolved mode at 77 K under 340 nm pulsed excitation after a 200 μ s delay between excitation and acquisition (Figure S11). The energy diagram reporting the ener-

gy positions measured for the S_1 and the T_1 states in **Gd-1** with respect to the electronic Ln^{3+} levels is presented on Figure 2.

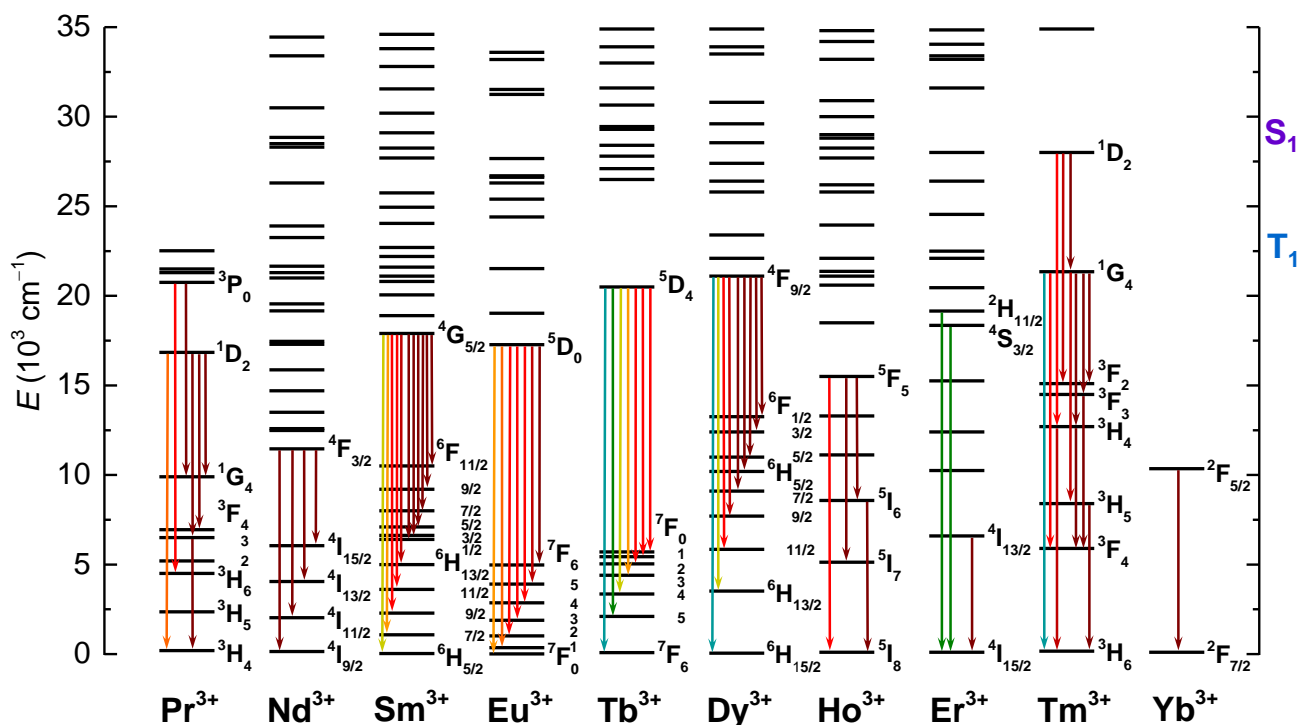


Figure 2. Lower energy part of the energy diagram of lanthanide(III) ions emitting in the visible and/or NIR range,⁷⁹⁻⁸¹ and energy positions of the ligands singlet (S_1) and triplet (T_1) states of **Ln-1** MCs. Radiative transitions observed for **Ln-1** MCs are indicated with arrows.

Lanthanide-centered photophysical properties

Excitation (Figure S12) and emission (Figure 3, S13) spectra, Ln^{3+} -centered quantum yields under ligand excitation (Q_{Ln}^{L}), and observed luminescence lifetimes (τ_{obs}) (Table 1) were collected for all **Ln-1** ($\text{Ln} = \text{Pr}^{3+}$, Nd^{3+} , Sm^{3+} , Eu^{3+} , Tb^{3+} – Yb^{3+}) in the solid state at room temperature. For the sake of comparison, Q_{Ln}^{L} values in Table 1 are reported separately for two ranges of emission wavelengths: 450–800 nm and 800–1650 nm. These are herein referred to as the visible and the NIR emission, respectively. In general, the emission efficiency of lanthanide(III) complexes can be described by the following equation:

$$Q_{\text{Ln}}^{\text{L}} = \eta_{\text{sens}} \cdot Q_{\text{Ln}}^{\text{Ln}} = \eta_{\text{sens}} \cdot \frac{\tau_{\text{obs}}}{\tau_{\text{rad}}} \quad (1)$$

where η_{sens} is the sensitization efficiency of the organic ligands that takes into account the energy migration pro-

cesses occurring both within the ligands and from the ligand to the Ln^{3+} excited state; $Q_{\text{Ln}}^{\text{Ln}}$ is the intrinsic quantum yield or quantum yield upon direct f-f excitation, it reflects the relative importance of radiative and non-radiative processes within Ln^{3+} ; τ_{rad} is the radiative lifetime, or the lifetime in absence of non-radiative deactivation processes. $Q_{\text{Ln}}^{\text{Ln}}$ is usually difficult to measure directly because of the very weak absorption of f-f transitions, however, it can be estimated as a ratio of τ_{rad} and the observed lifetimes (τ_{obs}).^{43, 82} The latter parameter is routinely accessed experimentally but the determination of τ_{rad} is less straightforward.⁸³ Presently, we estimated τ_{rad} for **Yb-1** using its absorption spectrum (Supporting Information, Figure S19), and for **Eu-1** using its emission spectrum (*vide infra*).

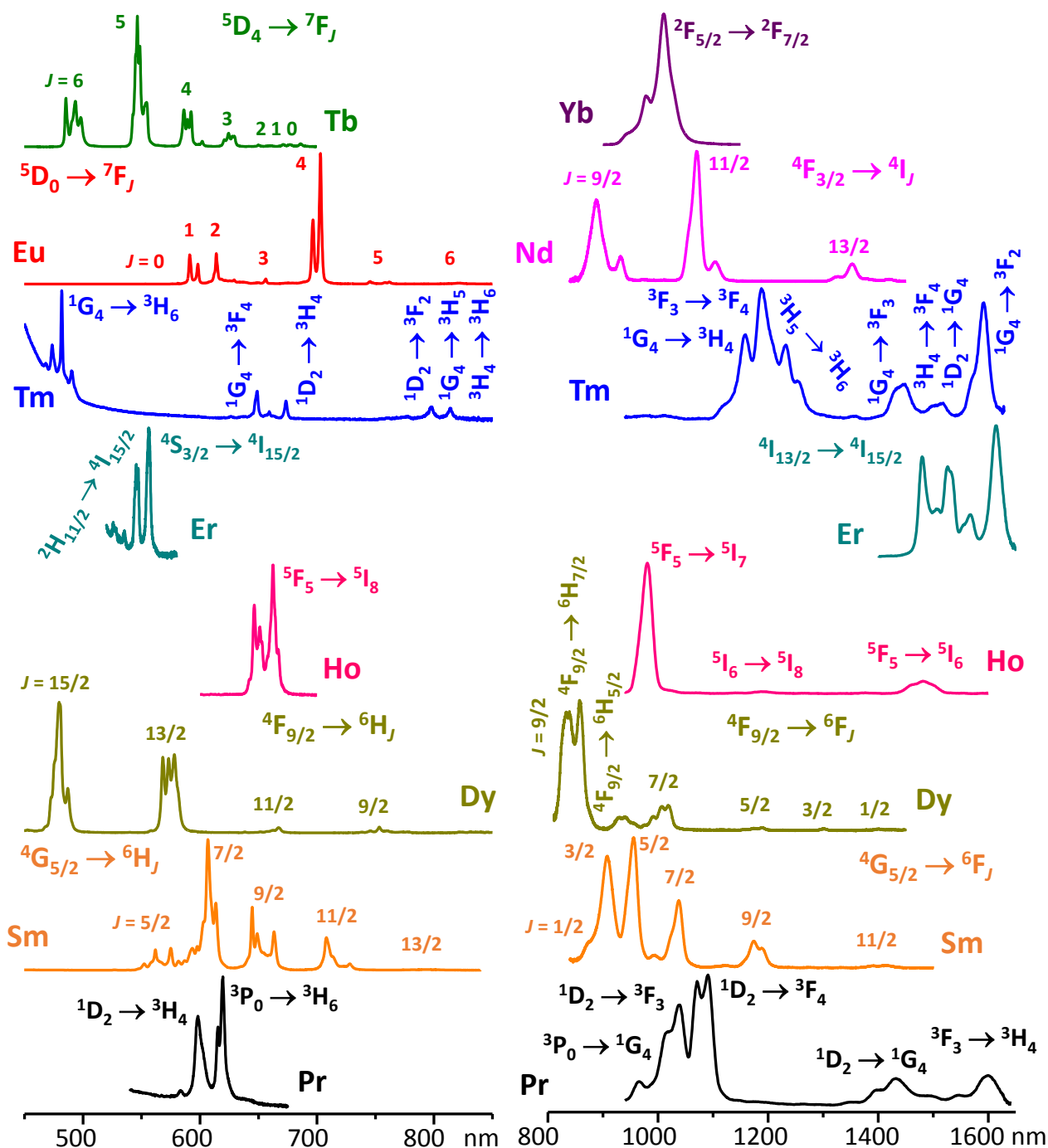


Figure 3. Highlights of the Ln³⁺-centered emission signals of Ln-1 MCs in the solid state under excitation at 340 nm at room temperature. Intensities of the emission spectra are normalized to the corresponding intensity maxima located in the range 450-850 nm (left) or 800-1650 nm (right).

To further establish the relationship between structure and luminescence properties, a high-resolution emission spectrum was acquired for **Eu-1** in the solid state (Figure S17, S18). To get information about the behavior of the studied MCs in solution and to assess the number of solvent molecules potentially coordinated to the Ln³⁺ (Supporting Information), a series of measurements was performed on solutions of selected Ln-1 (Ln = Nd³⁺, Sm³⁺, Eu³⁺, Tb³⁺, Dy³⁺, Yb³⁺) in CH₃OH and CD₃OD (Table 1, Figure 4,

S14, S15). For the less luminescent Ln-1 (Ln = Pr³⁺, Ho³⁺, Er³⁺, Tm³⁺), excitation and emission spectra were only collected in CD₃OD (Figure S16). The assignment of the bands in the excitation and emission spectra was done according to existing literature.^{79-81, 84-85}

Excitation spectra collected for Ln-1 in the solid state upon monitoring the main emission bands of the corresponding Ln³⁺ in the visible and/or the NIR ranges are dominated by broad ligand-centered bands in the UV up to

375 nm on the lower energy side, or 425 nm in the case of Eu-1 (Figure S12). The extension of the excitation band to the red for Eu-1 is in line with the diffuse reflectance spectrum (Figure S10, right) and the formation of LMCT state for this MC (*vide supra*) by describing the same electronic structure of the ligand-Eu³⁺ interaction. In addition to the broad bands present in the excitation spectra of all Ln-1 (except Yb-1) sharper features that correspond to f-f transitions that are specific to the nature of the specific Ln³⁺ are observed. These bands have different relative intensities compared to the ligand-centered ones and vary between the particular Ln³⁺ of interest. For example, direct f-f excitation bands are almost inconsequential in the case of Tb-1 while reaching 75% of the intensity of the ligand

maximum signal for Nd-1. Excitation spectra of Ln-1 measured in methanol solutions matches the absorption spectra and present only broad bands in the UV up to 350 nm. Expanded excitation range for solid state samples compared to those in solution can be explained by saturation effects.⁸⁶⁻⁸⁷ All these observations reflect the ability of the MC scaffold in Ln-1 to sensitize characteristic visible and NIR emissions of several Ln³⁺ of different natures through a common electronic structure.

Indeed, upon excitation into ligand-centered bands at 340 or 310 nm for samples in the solid state or in solution, respectively, Ln-1 exhibit sharp emission bands in the visible and/or NIR ranges that are specific to the corresponding f-f transitions (Figure 3, S13, S15, S16).

Table 1. Luminescence lifetimes (τ_{obs}) and Ln³⁺-centered quantum yields collected under ligand excitation (Q_{Ln}^{L}) for Ln-1.^a

Ln	State/ Solvent	τ_{obs} (μs) ^b	Level/ wavelength (nm) ^c	Q_{Ln}^{L} (%) ^d	
				450–800 nm	800–1650 nm
Pr	Solid	0.101(2)	¹ D ₂ / 1090	3.0(2)·10 ⁻²	4.91(6)·10 ⁻³
Nd	Solid	0.91(1)	⁴ F _{3/2} / 1070		0.30(2)
	CH ₃ OH	1.09(3) : 91(2)% ; 0.34(3) : 9(2)%	"		0.167(4)
	CD ₃ OD	5.9(1) : 89 (2)% ; 2.0(2) : 10(2)%	"		0.90(2)
Sm	Solid	151(1)	⁴ G _{5/2} / 611	10.4(6)	0.193(2)
	CH ₃ OH	83.4(3)	"	10.1(5)	7.98(3)·10 ⁻²
	CD ₃ OD	721(3)	"	83(1)	0.71(2)
Eu	Solid	1360(10)	⁵ D ₀ / 705	4.16(6)	
	CH ₃ OH	1950(1)	"	15.4(1)	
	CD ₃ OD	2500(10)	"	21(1)	
Tb	Solid	1040(50) : 87(1)% ; 390(30) : 13(1)%	⁵ D ₄ / 545	35(2)	
	CH ₃ OH	1870(10)	"	47.7(2)	
	CD ₃ OD	1950(10)	"	50.8(9)	
Dy	Solid	37.7(8) : 55(1)% ; 4.1(1) : 45(1)%	⁴ F _{9/2} / 577	1.26(7)	3.09(9)·10 ⁻²
	CH ₃ OH	24.8(2) : 87(1)% ; 2.69(9) : 13(1)%	"	6.9(3)	5.0(2)·10 ⁻²
	CD ₃ OD	41.5(8) : 92(1)% ; 4.54(7) : 8(1)%	"	9.46(6)	7.7(4)·10 ⁻²
Ho	Solid	0.0385(9)	⁵ F ₅ / 980	2.20(8)·10 ⁻²	2.9(3)·10 ⁻³
Er	Solid	1.93(6)	⁴ I _{13/2} / 1525	^e	1.90(4)·10 ⁻²
Tm	Solid	1.37(1)	¹ G ₄ / 482	5.8(2)·10 ⁻²	4.7(2)·10 ⁻⁵
Yb	Solid	27.4(9)	² F _{5/2} / 1010		2.8(1)
	CH ₃ OH	15.6(2)	"		0.78(4)
	CD ₃ OD	196(1)	"		8.4(1)

^a For samples in the solid state or for 50 μM solutions in CH₃OH or CD₃OD, at room temperature, 2 σ values between parentheses. Estimated experimental errors: τ_{obs} , $\pm 2\%$; Q_{Ln}^{L} , $\pm 10\%$. ^b Under excitation at 355 nm. If a biexponential decay was observed, popu-

lation parameters $P_i = \frac{B_i \tau_i}{\sum_{i=1}^n B_i \tau_i}$ in % are given after the colon. ^c Main emitting level and the corresponding wavelength used to record luminescence decays. ^d Under excitation at 340 nm for the samples in the solid state and at 310 nm for solutions. ^e Too low to be determined.

Pr-1 in the solid state exhibits both visible and NIR emission bands arising mainly from 1D_2 , but also from 3P_0 and 3F_3 excited states levels. Two visible emission bands are present in the range of 570–660 nm and are assigned to the $^1D_2 \rightarrow ^3H_4$ and $^3P_0 \rightarrow ^3H_6$ transitions with a Q_{Pr}^L value of $3.0 \cdot 10^{-2}\%$. In the NIR domain, emission bands arise from the $^3P_0 \rightarrow ^1G_4$, $^1D_2 \rightarrow ^3F_3$, 3F_4 , 1G_4 , and $^3F_3 \rightarrow ^3H_4$ electronic transitions and are observed in the range 950–1650 nm. Q_{Pr}^L (NIR) was found to be $4.9 \cdot 10^{-3}\%$, while the τ_{obs} value of the 1D_2 level is 101(2) ns. NIR emission of weaker intensity of **Pr-1** could be detected in CD₃OD solution (Figure S16).

Nd-1 in the solid state, solutions in CH₃OH and CD₃OD shows NIR luminescence in the range of 850–1450 nm with three main bands at ~885, 1065 and 1350 nm due to the $^4F_{3/2} \rightarrow ^4I_J$ ($J = 9/2-13/2$) transitions. The general features of each of them remain similar for the sample in the solid state and in the solution. However, the relative intensity of the emission band corresponding to the $^4F_{3/2} \rightarrow ^4I_{9/2}$ transition significantly increases for the latter. Luminescence decays are monoexponential for the solid state sample and are biexponential in solution, with the 10% contribution of a short-lived component. τ_{obs} values are comparable for the solid state sample and in solution in CH₃OH while Q_{Nd}^L is decreased from 0.3 to 0.17%. In CD₃OD solution, both τ_{obs} and Q_{Nd}^L are significantly increased reaching values of 5.9 μ s and 0.9%, respectively.

Sm-1 emits in both visible and NIR, emission signals arising from the $^4G_{5/2}$ level in the solid state and solutions. Multiple sharp emission bands in the visible range can be assigned to $^4G_{5/2} \rightarrow ^3H_J$ ($J = 5/2-13/2$) transitions with $^4G_{5/2} \rightarrow ^3H_{7/2}$ dominating the spectrum and being responsible for the pink-orange color of Sm-1 emission. NIR emission is derived from $^4G_{5/2} \rightarrow ^6F_J$ ($J = 5/2-13/2$) transitions. The crystal-field splitting of the emission bands is very similar for the **Sm-1** in the solid state and in solution but the relative intensities of the $^4G_{5/2} \rightarrow ^3H_{9/2, 11/2}$, $^6F_{7/2, 9/2}$ transitions are smaller for the latter. A residual broad-band ligand-centered emission is observed at wavelengths <450 nm. Quantitative data are provided in Table S6. The values of Q_{Sm}^L (visible) are essentially the same (~10%) for the sample in the solid state and in the CH₃OH solution, while τ_{obs} are 1.8-times shorter for the latter. Both parameters are drastically enhanced in CD₃OD solution reaching 83% and 721 μ s. The values of Q_{Sm}^L in the NIR are lower and increase in the following order: CH₃OH < solid state < CD₃OD.

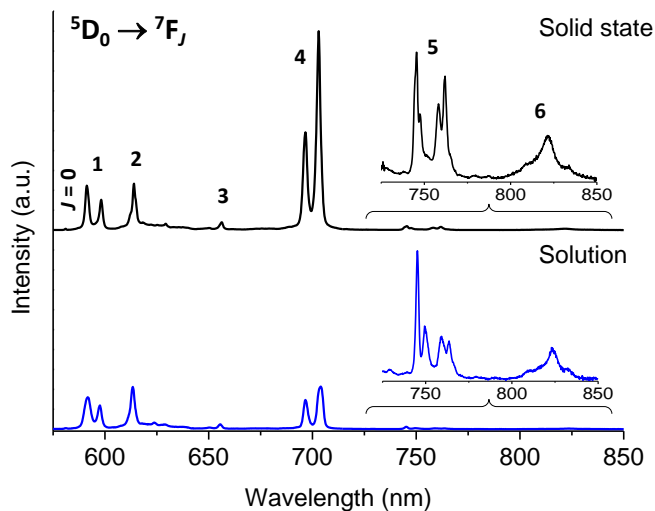


Figure 4. Comparison between emission spectra for **Eu-1** in the solid state ($\lambda_{ex} = 340$ nm) and in the 50 μ M methanol solution ($\lambda_{ex} = 310$ nm) at room temperature (emission bandpass 1 nm, acquisition increment 0.2 nm). The spectra have been scaled so that the respective $^5D_0 \rightarrow ^7F_1$ transitions have identical integrated areas.

Eu-1 shows red luminescence with sharp emission bands assigned to $^5D_0 \rightarrow ^7F_J$ ($J = 0-6$) transitions (Figure 4). The strictly forbidden $^5D_0 \rightarrow ^7F_0$ transition results in an emission band possessing very low intensity. However, all the other transitions including the rarely observed $^5D_0 \rightarrow ^7F_{5,6}$ can be clearly distinguished. In the solid state, the **Eu-1** emission spectrum is dominated by the $^5D_0 \rightarrow ^7F_4$ transition that has 60.9% contribution to the total emission intensity (Table S4). The hypersensitive $^5D_0 \rightarrow ^7F_2$ and magnetic dipole $^5D_0 \rightarrow ^7F_1$ transitions have almost equal contribution with the ratio $I(^5D_0 \rightarrow ^7F_2)/I(^5D_0 \rightarrow ^7F_1)$ equal to 1.25. Such observations are consistent with a square anti-prismatic environment around Eu³⁺ and a pseudo- D_{4h} symmetry (*vide supra*).⁸⁸ At room temperature Q_{Eu}^L is equal to 4.2 % while τ_{obs} is almost temperature independent (1.36 ms at 298K vs. 1.55 ms 10K).

To gain more insight about Eu³⁺ environment, a high-resolution emission spectrum of **Eu-1** in the solid state was recorded at 10K (Figure S18). The narrow $^5D_0 \rightarrow ^7F_0$ transition has a full-width at half-maximum (FWHM) of 10 cm^{-1} and a maximum at 17 209 cm^{-1} (Figure S17). The observation of a single emission band for this transition reflects the presence of only one type of emitting Eu³⁺ species in **Eu-1** and therefore of a unique environment around this metal ion. In addition, it is generally accepted that the energy of the $^5D_0 \rightarrow ^7F_0$ transition ($\tilde{\nu}$) and, in particular, its shift from the energy determined for Eu³⁺ in the gas phase ($\tilde{\nu}_0 = 17 374$ cm^{-1}) depends on the nephelauxetic effect generated by the coordinating ligands. Several ap-

proaches have been suggested to estimate this effect,⁸⁹ but if considering the one by Frey and Horrocks:⁹⁰

$$\tilde{\nu} = \tilde{\nu}_0 + C_{CN} \sum_{i=1}^{CN} n_i \delta_i \quad (2)$$

where δ_i is the nephelauxetic parameter for each coordinating ligand or group, n_i is the index of such groups, CN is the coordination number (CN = 8 for **Eu-1**), and C_{CN} is a coefficient that depends on CN ($C_{CN} = 1.06$ for CN = 8). If taking into account an approximation that $\tilde{\nu}$ increase by 1 cm^{-1} when the temperature rises by 24 K, $\tilde{\nu}$ equals to 17 221 cm^{-1} at 298K. Therefore, $\delta_i = -18$ for the shi^{3-} ligand's binding hydroximate oxygens, a value which is close to the effect induced by a charged carboxylate oxygen ($\delta_i = -18.2$).⁹⁰ The ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition contains two-fold degenerate and non-degenerate components that correspond to ${}^7\text{F}_1$ crystal-field sublevels located at 285.1, 335.3 and 506.2 cm^{-1} above the ${}^7\text{F}_0$ level. Such splitting reflects a pseudo-tetragonal symmetry around Eu^{3+} consistent with the square antiprism coordination environment.

The crystal-field splitting of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_j$ ($J = 0-6$) transitions in the emission spectrum of **Eu-1** in methanol solution are very similar to those observed in the solid state (Figure 4). However, the contribution of the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ transition to the total emission intensity is decreased to 34.1% (vs. 60.9% in the solid state). On the other hand, the ratio $I({}^5\text{D}_0 \rightarrow {}^7\text{F}_2)/I({}^5\text{D}_0 \rightarrow {}^7\text{F}_1)$ remains comparable (1.18 vs. 1.25) (Table S4). Compared to the solid state sample, the values of τ_{obs} and Q_{Eu}^{L} increase by 1.4 and 3.7 times, respectively, for **Eu-1** in CH_3OH solution. In deuterated solvent, these parameters increase by 1.8 and 5.0 fold respectively (Table 2).

To get more information about the parameters that affect the **Eu-1** emission intensity, τ_{rad} and the corresponding $Q_{\text{Eu}}^{\text{Eu}}$ and η_{sens} were estimated using eq. 1 (Table 2). In the case of Eu^{3+} , the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition has a purely magnetic dipolar character, is insensitive to the Eu^{3+} coordination environment and, when J -mixing is considered to be small, the integrated intensity of the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition (I_{MD}) can be used to calibrate the total integrated emission intensity arising from ${}^5\text{D}_0 \rightarrow {}^7\text{F}_j$ ($J = 0-6$) transitions (I_{tot}) and to determine τ_{rad} via the equation:^{82,91}

$$\frac{1}{\tau_{\text{rad}}} = A_{\text{MD},0} \cdot n^3 \left(\frac{I_{\text{tot}}}{I_{\text{MD}}} \right) \quad (3)$$

where $A_{\text{MD},0}$ is the spontaneous emission probability of the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition being equal to 14.65 s^{-1} , n is the refractive index. τ_{rad} are 2.8 fold longer for **Eu-1** in solution compared to those recorded in the solid state, leading to a 1.5-2 times decrease of the intrinsic quantum yields. On the other hand, the 7.3 fold increased sensitization efficiency of the ligands in solutions results in an overall increase of Q_{Eu}^{L} .

Table 2. Photophysical parameters obtained for **Eu-1 in the solid state and in 50 μM solutions.^a**

State/ Solvent	τ_{obs} (ms)	τ_{rad} (ms) ^b	$Q_{\text{Eu}}^{\text{Eu}}$ (%)	Q_{Eu}^{L} (%)	η_{sens} (%)
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Solid state	1.36(1)	2.78	49	4.16(6)	8.5
CH_3OH	1.95(1)	7.80	24	15.4(5)	62
CD_3OD	2.50(1)	7.87	32	21(1)	66

^a At room temperature, 2σ values between parentheses. Relative errors: τ_{obs} , $\pm 2\%$; Q_{Eu}^{L} , $\pm 10\%$; τ_{rad} , $\pm 10\%$; $Q_{\text{Eu}}^{\text{Eu}}$, $\pm 12\%$; η_{sens} , $\pm 22\%$. ^b Calculated using Eq. 3 according to ref. ⁹¹; $n(\text{solid state}) = 1.5$; $n(\text{CH}_3\text{OH}) = 1.33$; $n(\text{CD}_3\text{OD}) = 1.326$.

Tb-1 in the solid state and in solution exhibits a bright green emission arising from ${}^5\text{D}_4 \rightarrow {}^7\text{F}_j$ ($J = 6-0$) transitions with Q_{Tb}^{L} of 35-51%. The crystal field splitting of the ${}^5\text{D}_4 \rightarrow {}^7\text{F}_j$ bands is similar for **Tb-1** in the solid state and in solution (Figure S14). On the other hand, the relative integral intensity of the ${}^5\text{D}_4 \rightarrow {}^7\text{F}_6$ transition increases by 1.5 fold while for the other ${}^5\text{D}_4 \rightarrow {}^7\text{F}_j$ ($J = 4-0$) transitions a diminution by 1.4-3 fold is observed (Table S5). Experimental luminescence decays arising from the ${}^5\text{D}_4$ level in the solid state are best fitted with a biexponential function with the dominating contribution (87 %) from the long component (1040 μs) and an average τ_{obs} value of 1010 μs . τ_{obs} measured at 10 K were found to be $\tau_1 = 1040(10)$ μs ($P_1 = 96.4(4)$ %) and $\tau_2 = 272(5)$ μs ($P_2 = 3.6(4)$ %) giving an average value of 1030 μs . The similarity of the average values of τ_{obs} when going from 298 K to 10 K reflects the absence or an insignificant role of temperature dependent non-radiative processes including Boltzmann dependent back energy transfer in **Tb-1**. Luminescence decays of **Tb-1** in solution are monoexponential and the corresponding τ_{obs} values are ~ 2.2 times longer than those recorded in the solid state.

Dy-1 in the solid state and in solution emits both visible and NIR light originating from the ${}^4\text{F}_{9/2}$ level.⁸⁵ In the visible regime, the emission spectrum is dominated by the ${}^4\text{F}_{9/2} \rightarrow {}^6\text{H}_{15/2}$ transition centered at 480 nm, while in the NIR region the emission is dominated by the bands associated with the ${}^4\text{F}_{9/2} \rightarrow {}^6\text{H}_{9/2, 7/2}$ transitions in the range of 800-880 nm. A residual broad-band ligand-centered emission <450 nm is observed (Figure S14, Table S6). The values of Q_{Dy}^{L} (visible) of **Dy-1** in CH_3OH and CD_3OD are 5.5 and 7.5 times higher than in the solid state. Dy^{3+} -centered emission in the NIR is also enhanced in solution compared to the solid state but to a lesser extent, by a 1.6-2.5 fold. Luminescence decays curves recorded at 298 K upon monitoring the emission from the ${}^4\text{F}_{9/2}$ level are biexponential with an almost equal contribution from the short and the long components, and the average value of 35 μs . At 10 K, the latter value is 3.5 times longer (123 μs) in respect to $\tau_1 = 125(1)$ μs ($P_1 = 92.7(6)\%$) and $\tau_2 = 35(1)$ μs ($P_2 = 7.3(6)\%$). Such behavior can be explained by the existence of the temperature-dependent non-radiative energy transfer processes such as the back transfer from the $\text{Dy}^{3+} {}^4\text{F}_{9/2}$ level ($21\,100\text{ cm}^{-1}$)⁸¹ to the closely located triplet state of the ligands ($22\,620\text{ cm}^{-1}$). In contrast to the behavior of **Tb-1**, **Dy-1** luminescence decays in solutions remain biexponential, although the contribution of the longer component is increased to 87-92 % vs. 55 % for the solid state. The average τ_{obs} values are similar for the solid state

sample and the solution in CH₃OH (19.1 μs and 21.1 μs) but they increase to 37.9 μs for the CD₃OD solution.

Ho-1 in the solid state exhibits both visible and NIR emissions. The visible emission in the range of 635–680 nm is attributed to the ⁵F₅→⁵I₈ transition and has a Q_{Ho}^{L} (visible) of 2.20(8)·10⁻² %. NIR emission signal is dominated by the ⁵F₅→⁵I₇ transition at 980 nm with smaller contributions from the ⁵F₅→⁵I₆ band at 1480 nm and a faint participation of the ⁵I₆→⁵I₈ transition at 1190 nm. The observed luminescence lifetime of the ⁵F₅ level is 38.5(9) ns.

Er-1 in the solid state shows multiple emission bands located in the range 1450–1650 nm and assigned to the ⁴I_{13/2}→⁴I_{15/2} transition. The luminescence decay of the ⁴I_{13/2} level is monoexponential with a τ_{obs} value of 1.93(6) μs. Q_{Er}^{L} (NIR) is equal to 1.90(4)·10⁻² %. In addition to the NIR emission, faint Er³⁺ transitions in the visible range between 520 and 565 nm due to ²H_{11/2}→⁴I_{15/2} and ⁴S_{3/2}→⁴I_{15/2} transitions were detected, however quantitative parameters related to these bands could not be determined.

Tm-1 possesses faint visible and NIR emissions with multiple bands arising from the ¹G₄, ¹D₂, ³H₄ and ³F₃ excited levels.⁸⁴ In the visible range, the ¹G₄→³H₆ transition with a maximum at 480 nm overlaps with the broad ligand-centered emission (Figure S13). Other bands are observed in the range of 620–840 nm. The total Q_{Tm}^{L} (visible) equal to 5.8(2)·10⁻² % and the τ_{obs} of the ¹G₄ level is 1.37(1) μs. Moreover, Tm³⁺ f-f specific electronic transitions with maxima at 1188, 1448 and 1592 nm could be detected in the range of 1100–1650 nm with Q_{Tm}^{L} (NIR) of 4.7(2) 10⁻⁵ %.

Yb-1 in the solid state and solutions displays NIR emission in the range of 920–1100 nm with a maximum at 1010 nm arising from the ²F_{5/2}→²F_{7/2} transition. The values of τ_{obs} and Q_{Yb}^{L} decrease by 1.8 and 3.6 times, respectively, when comparing **Yb-1** in the solid state and in CH₃OH solution. In deuterated methanol, both values are significantly enhanced, reaching 196 μs and 8.4 %. To get additional information about the different parameters that impact the global **Yb-1** emission efficiency, τ_{rad} was estimated from the absorption spectrum measured in the range of the ²F_{5/2}←²F_{7/2} transition (Figure S19) using the modified Einstein’s equation (Supporting Information).⁹¹ The data are provided in Table 3 together with the values reported previously for other Yb³⁺/Ga³⁺ MCs for comparison.

Table 3. Photophysical parameters obtained for Yb-1 in solution and comparison with other Yb^{III}/Ga^{III} MCs.^a

MC	Solvent	τ_{obs} (μs)	τ_{rad} (μs) ^b	$Q_{\text{Yb}}^{\text{Yb}}$ (%)	Q_{Yb}^{L} (%)	η_{sens} (%)
Yb-1	CH ₃ OH	15.6(2)	814	1.9	0.78(4)	41
	CD ₃ OD	196(1)	819	24	8.4(1)	35
YbGa ₄ (1) ^c	CH ₃ OH	2.06(4)	530	0.39	0.26(1)	67

	CD ₃ OD	36.6(1)	530	6.9	4.29(1)	62
YbGa ₄ (2) ^d	CH ₃ OH	1.72(1)	590	0.29	0.123(2)	42
	CD ₃ OD	27.4(1)	590	4.6	2.55(5)	55

^a At room temperature, 2σ values between parentheses. Relative errors: τ_{obs} , ±2%; Q_{Yb}^{L} , ±10%; τ_{rad} , ±10%; $Q_{\text{Yb}}^{\text{Yb}}$, ±12%; η_{sens} , ±22%. ^b See Supporting Information for details; for CD₃OD solutions recalculated taking into account the refractive index $n(\text{CD}_3\text{OD}) = 1.326$. ^c [YbGa₄(shi)₄(benzoate)₄] from Ref.⁵ ^d [YbGa₄(shi)₄(H₂shi)₂] from Ref.⁶

DISCUSSION

Synthesis and characterization

Ln-1 is the direct Ln³⁺-centered analogue of the [Na₃Ga₈(shi)₈(OH)₄] structure published previously by our group (Figure 1a).⁴ It is also a major evolution of the Ga³⁺/Ln³⁺ luminescent MCs created by us initially with the monomeric [LnGa₄(shi)₄(benzoate)₄] (Figure 1b),⁵ then with the [Ln₂Ga₈(shi)₈(isophthalate)₄] (Figure 1c)³ series. The [LnGa₈(shi)₈(OH)₄] MC (Figure 1d) possesses a pseudo-*D*_{4h} symmetry about the central Ln³⁺ as compared to the previously mentioned Ga³⁺/Ln³⁺ MCs in which the Ln³⁺ is located on a pseudo-*C*₄ axis of symmetry.

The ESI-MS mass-spectra of **Ln-1** measured in acetonitrile and methanol solutions revealed the exclusive presence of the MC species. The DOSY NMR data performed on **Y-1** further supports this observation (Supporting Information). Together, these results suggest that the **Ln-1** MCs are present as stable dimers in solution. However, in methanol solution, as indicated by mass spectrometry experiments, a rapid OH⁻/CH₃O⁻ exchange occurs which will eventually lead to CH₃O⁻ becoming the dominant bridging ion and [LnGa₈(shi)₈(CH₃O)₄] becoming the main species.

Photophysical properties

In the solid state, upon excitation into the ligand-centered bands at 340 nm, **Eu-1** and **Tb-1** exhibit bright red and green emission signals while **Yb-1** and **Nd-1** show their respective characteristic transitions in the NIR range (Figure 3). For **Pr-1**, **Sm-1**, **Dy-1**, **Ho-1**, **Er-1** and **Tm-1** dual-range emissions with bands located in the visible and in the NIR ranges were detected. It should be noted such observation of dual emission for six Ln³⁺ using one type of molecular compound are very scarce.^{47, 50, 56, 62, 92} Additionally, if the dual emission of Sm³⁺ and Dy³⁺ are more frequently observed and quantified,^{3, 5-6, 46-47, 57, 60, 65, 67-68, 85, 93-94} studies that involve evaluation of such properties from Pr³⁺, Ho³⁺, Er³⁺ and Tm³⁺ in coordination compounds are very rare.^{58-59, 61, 64, 95} However, unique bands arising from Ln³⁺ of different natures and spanning the whole visible and NIR ranges with various luminescence lifetimes and emission profiles are essential for multicolor and multiplex imaging,⁹⁶⁻⁹⁷ and could lead to the development of advanced materials such as sophisticated barcode modules⁶³⁻⁶⁵ or other interesting optical imaging applications.⁶⁷⁻⁶⁸

It is generally accepted that the lowest in energy triplet state of chromophoric ligands play a predominant role in the sensitization of Ln^{3+} although singlet and charge transfer (CT) states may also be involved.^{30, 43} The energy position of the T_1 level determined from the phosphorescence spectrum of **Gd-1** was found at 22 620 cm^{-1} , a value similar to the ones observed for the previously reported series of MCs possessing $\text{Ga}^{3+}/\text{shi}$ scaffold, i.e. $[\text{LnGa}_4(\text{shi})_4(\text{H}_2\text{shi})_2]$ (21 660 cm^{-1}),⁶ $[\text{LnGa}_4(\text{shi})_4(\text{benzoate})_4]$ (22 170 cm^{-1})⁵ or $[\text{Ln}_2\text{Ga}_8(\text{shi})_8(\text{isophthalate})_4]$ (21 190 cm^{-1}).³ The T_1 level in **Ln-1** is located slightly higher in energy than the emitting level of Dy^{3+} , $\Delta E (T_1 \leftarrow {}^4F_{9/2}) = 1\,520\text{ cm}^{-1}$, and a bit further from that of Tb^{3+} , $\Delta E (T_1 \leftarrow {}^5D_4) = 2\,220\text{ cm}^{-1}$ (Figure 2). Relatively small difference in energy between the two levels increase the probability of thermally activated back $T_1 \leftarrow \text{Ln}^{3+*}$ energy transfer. A clear influence of this process was confirmed for **Dy-1** with the 3.5 times lengthening of the average τ_{obs} value upon cooling from room temperature to 10 K. In the case of **Tb-1**, this effect is insignificant and averaged τ_{obs} values are similar at both temperatures. Back energy transfer is probably responsible for the bi-exponential character of luminescence decays observed for **Tb-1** and **Dy-1** (Table 1). For **Tm-1** in the solid state, emission is originating from different electronic levels (Figures 2, 3), i.e. 1D_2 (28 000 cm^{-1}), 1G_4 (21 350 cm^{-1}), 3H_4 (12 700 cm^{-1}), and 3H_5 (8 400 cm^{-1}). The T_1 level is positioned well above the ${}^3H_{4,5}$ levels but close to 1G_4 ($\Delta E (T_1 \leftarrow {}^1G_4) = 1\,270\text{ cm}^{-1}$) and is significantly lower in energy than the 1D_2 level ($\Delta E (T_1 \leftarrow {}^1D_2) = -5\,380\text{ cm}^{-1}$). The latter point suggests that Tm^{3+} -centered luminescence in **Tm-1** could be partially sensitized through the lowest singlet state of the ligands located at 28 690 cm^{-1} . For **Pr-1**, most of the emission bands originate from the 1D_2 (16 840 cm^{-1}) excited level, however transitions from the 3P_0 (20 750 cm^{-1}) in the visible and the NIR ranges and from the 3F_3 (6 500 cm^{-1}) in the NIR could also be detected (Figures 2, 3). Emitting levels of other studied Ln^{3+} are located at least 3 470 cm^{-1} below the energy of the T_1 level making it potentially well-adapted for their sensitization (Figure 2).

In the case of **Eu-1** in the solid state, the slight broadening to the red of the diffuse reflectance spectrum is an indication of the presence of a LMCT. Such electronic states, depending on their energy position, can either quench or enhance Eu^{3+} luminescence.^{45, 49, 98} Contrary to all of the $\text{Ga}^{3+}/\text{shi}$ systems previously described by us where the formation of LMCT states either significantly⁵ or fully^{3, 6, 70} quench Eu^{3+} emission, **Eu-1** exhibited pronounced red emission (Figure 3). Therefore, this series of $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]$ MCs is the first $\text{Ga}^{3+}/\text{shi}$ MC scaffold that can efficiently sensitize Eu^{3+} luminescence. This luminescence originating mainly from ${}^5D_0 \rightarrow {}^7F_j$ ($j = 0-6$) transitions is not only interesting for numerous practical applications⁵² but also from a theoretical point of view.⁸⁸ The non-degenerate character of the emitting 5D_0 level allows the easier interpretation of Eu^{3+} luminescence spectra. Therefore, the analysis of the relative intensities of ${}^5D_0 \rightarrow {}^7F_j$ ($j = 0-6$) transitions and their crystal-field splitting provide information about the point group symmetry and the coordination polyhedron around Eu^{3+} . Thus, interestingly, the emission spectrum of **Eu-1** in the solid state is domi-

nated by the ${}^5D_0 \rightarrow {}^7F_4$ transition while the hypersensitive ${}^5D_0 \rightarrow {}^7F_2$ and magnetic dipole ${}^5D_0 \rightarrow {}^7F_1$ transitions have almost equal contributions (Figure 4, Table S4). These features reflect the square anti-prismatic environment around the Eu^{3+} and the presence of a pseudo- D_{4h} symmetry,⁸⁸ in agreement with structural data. Further confirmation about the symmetry was obtained through the analysis of the fine splitting of the ${}^5D_0 \rightarrow {}^7F_1$ transition into two-fold degenerate and non-degenerate components in the high-resolution spectrum (Figure S17). Moreover, the presence of a single narrow ${}^5D_0 \rightarrow {}^7F_0$ transition (FWHM = 10 cm^{-1}) in the latter reflects the presence of only one type of emitting Eu^{3+} and therefore a unique and well-defined coordination environment around Eu^{3+} in **Eu-1**.

Dual emission of **Er-1** deserves a special attention. Multiple emitting levels and small energy gaps facilitating non-radiative deactivation processes make the design of luminescent Er^{3+} compounds with organic ligands highly challenging.³⁰ Nevertheless, potential applications in telecommunications^{34, 99} of the $\text{Er}^{3+} {}^4I_{13/2} \rightarrow {}^4I_{15/2}$ transition with an emission wavelength at $\sim 1500\text{ nm}$ have initiated multiple attempts to create and enhance Er^{3+} emission in coordination compounds.¹⁰⁰⁻¹⁰² The observation of Er^{3+} transitions in the visible range is uncommon and only a few examples could be found in the literature.^{56, 58, 103-104} Nevertheless, we could unambiguously detect ${}^2H_{11/2} \rightarrow {}^4I_{15/2}$ and ${}^4S_{3/2} \rightarrow {}^4I_{15/2}$ transitions between 520 and 565 nm and as multiple bands in the range of 1450-1650 nm due to the ${}^4I_{13/2} \rightarrow {}^4I_{15/2}$ transition in the emission spectrum of **Er-1** in the solid state (Figure 2).

In CH_3OH and CD_3OD , solutions of **Ln-1** ($\text{Ln} = \text{Nd}^{3+}, \text{Sm}^{3+}, \text{Eu}^{3+}, \text{Tb}^{3+}, \text{Dy}^{3+}, \text{Yb}^{3+}$) exhibit characteristic f-f transitions in the visible and/or NIR ranges upon excitation on chromophoric ligands at 310 nm. Compared to the corresponding emission spectra recorded in the solid state, a general trend was detected for all studied **Ln-1** in solution: the crystal-field splitting of f-f transitions remains the same while their relative intensities vary. The most significant change was observed for the **Eu-1** emission spectrum, which in the solid state is dominated by the ${}^5D_0 \rightarrow {}^7F_4$ transition. However, in solution, this band has an equal intensity to the hypersensitive ${}^5D_0 \rightarrow {}^7F_2$ transition (Figure 4, Table S4). Such variations are usually caused by changes in coordination environment and symmetry around Ln^{3+} .⁸⁸ However, NMR studies and ESI mass-spectrometry (*vide supra*) showed that the MC structure remains intact in solution while four OH^- bridges can be exchanged with $\text{CH}_3\text{O}^-/\text{CD}_3\text{O}^-$. Additionally, ESI mass spectrometry data show that, in solution, the solvation of the complex ion is likely to remove the coordination of two (1/2 equivalence) Na^+ ions bound at one side of the complex. This situation may lead to a more symmetrical structure in solution than in the solid state. Such behavior can affect the global electronic structure of the MC scaffold and may explain the changes of relative intensities of f-f transitions. Phenomenological equations¹⁰⁵ were used to estimate the number of solvent molecules coordinated to the Ln^{3+} based on the difference in τ_{obs} values recorded in both CH_3OH and CD_3OD solutions for **Eu-1**, **Tb-1**, and **Yb-1**. Results con-

firmed that this MC scaffold protects the coordinated Ln^{3+} ion well and no solvent molecules are directly coordinated. It should be noted, however, that in the case of **Nd-1**, luminescence decays become bi-exponential in solution possibly because of the partial coordination of solvent molecules. Such observation could be explained by Nd^{3+} having a larger ionic radius, providing additional space for the access of quenching solvent molecules.

Quantitative photophysical parameters, in particular the values of τ_{obs} and Q_{Ln}^{L} for **Ln-1** in the solid state, are comparable or superior (for **Eu-1** and **Sm-1**) to those reported previously for the $[\text{LnGa}_4(\text{shi})_4(\text{benzoate})_4]^5$ or $[\text{Ln}_2\text{Ga}_8(\text{shi})_8(\text{isophthalate})_4]^3$ MCs (Table S7). In protic solvents, the values of τ_{obs} and Q_{Ln}^{L} are usually lower than those measured in the solid state because of the higher probability of non-radiative deactivation through overtones of O–H vibrations.³⁰ Indeed, the values of Q_{Ln}^{L} measured for **Nd-1**, **Yb-1** and the NIR ($^4\text{G}_{5/2} \rightarrow ^6\text{F}_j$) transitions of **Sm-1** are smaller in CH_3OH than in the solid state (Table 1). Nevertheless, the changes caused by the dissolution in a protic solvent are less significant than those reported previously for $[\text{LnGa}_4(\text{shi})_4(\text{benzoate})_4]^5$. Moreover, in the case of **Eu-1**, **Tb-1**, **Dy-1** or visible ($^4\text{G}_{5/2} \rightarrow ^3\text{H}_j$) transitions of **Sm-1** the values of Q_{Ln}^{L} are, respectively, 3.7-, 1.4-, 5.5-times higher or the same in CH_3OH solution than in the solid state. It should be noted that Q_{Sm}^{L} of **Sm-1** in CH_3OH solution ($\sim 10\%$) is exceptionally high since this value usually does not exceed 2% for Sm^{3+} complexes in protic solvents,¹⁰⁶ while Q_{Yb}^{L} of **Yb-1** is 3-times larger in respect to the value recorded for $[\text{YbGa}_4(\text{shi})_4(\text{benzoate})_4]^5$ (Table 3). In deuterated CD_3OD solutions of **Ln-1**, the values of τ_{obs} and Q_{Ln}^{L} are enhanced compared to those obtained in CH_3OH being in line with lower contribution of O–D vibrations to the non-radiative deactivation of Ln^{3+} .¹⁰⁵ This effect is the most sizable for **Sm-1** and **Yb-1**, the τ_{obs} and Q_{Ln}^{L} of which increase by 8–12 times reaching values of 721 $\mu\text{s}/83\%$ and 196 $\mu\text{s}/8.4\%$, respectively (Table 1). These values are exceptional for Sm^{3+} coordination compounds^{59, 106–108} and are among the highest reported today for Yb^{3+} complexes formed with non-deuterated and non-halogenated ligands.^{51, 69, 109–111}

To quantify the parameters that impact the total quantum yield (Q_{Ln}^{L}) and to get a potential explanation for the photophysical properties of the studied series of $\text{Ln}^{3+}/\text{Ga}^{3+}$ MCs, the values of τ_{rad} were determined for **Eu-1** (Table 2) and **Yb-1** (Table 3) using Eqs. (3) and (4a,b), respectively. Further, knowing τ_{rad} , τ_{obs} and Q_{Ln}^{L} (Table 1), it is possible to estimate intrinsic quantum yields ($Q_{\text{Ln}}^{\text{In}}$) and sensitization efficiencies of the ligands, η_{sens} , according to Eq. (1).

It was found that τ_{rad} is equal to 2.78 ms for **Eu-1** in the solid state, a value that is close to the one reported, for example, for Eu^{3+} complexes formed with benzothiazole-substituted pyridine-2-carboxylate ligands (2.7–3.2 ms).¹¹² In CH_3OH or CD_3OD solutions, the value of τ_{rad} becomes 2.8-times longer (7.8 ms). Similar values were reported for aqueous solutions of Eu^{3+} helicates with pseudo- D_3 symmetry (6.2–6.9 ms).¹¹³ Such changes of τ_{rad} are most probably induced by an exchange of four OH^- bridges with $\text{CH}_3\text{O}^-/\text{CD}_3\text{O}^-$ in solution as suggested by mass-

spectrometry data (*vide supra*). The lengthening of τ_{rad} leads to a decrease of $Q_{\text{Eu}}^{\text{Eu}}$ by 1.5–2 times in solution compared to the value obtained in the solid state. Nevertheless, significantly improved η_{sens} (7.3–7.6 fold) in solutions of **Eu-1** results in a 3.7–5 times enhancement of Q_{Eu}^{L} .

In the case of **Yb-1**, the different photophysical parameters obtained in CH_3OH or CD_3OD solutions could be compared with the values reported for $[\text{YbGa}_4(\text{shi})_4(\text{benzoate})_4]^5$ and $[\text{YbGa}_4(\text{shi})_4(\text{H}_2\text{shi})_2]^6$ (Table 3). The values of τ_{rad} for **Yb-1** are 1.4–1.5 times longer than for the other $\text{Yb}^{3+}/\text{Ga}^{3+}$ MCs, probably, due to the higher level of symmetry around Yb^{3+} for the former one (pseudo- D_{4h} vs. pseudo- C_4 or C_1). However, the values of $Q_{\text{Yb}}^{\text{Yb}}$ are significantly (4.9–6.6 fold) higher for **Yb-1** than for the $[\text{YbGa}_4(\text{shi})_4(\text{benzoate})_4]$ and $[\text{YbGa}_4(\text{shi})_4(\text{H}_2\text{shi})_2]$ MCs because of the longer τ_{obs} . Notably, these two parameters observed for **Yb-1** are similar to the ones reported for $\text{Yb}^{3+}/\text{Zn}^{2+}$ “encapsulated sandwich” MCs with quinaldichydroxamic acid, which similarly have a D_{4h} type symmetry.¹¹⁴ Finally, sensitization efficiencies for **Yb-1** are comparable or slightly lower than in $[\text{YbGa}_4(\text{shi})_4(\text{benzoate})_4]$ and $[\text{YbGa}_4(\text{shi})_4(\text{H}_2\text{shi})_2]$. Together with improved $Q_{\text{Yb}}^{\text{Yb}}$ and τ_{obs} for **Yb-1**, we observed a more than 3.5–6.5 times enhancement of Q_{Yb}^{L} values in solution relative to the aforementioned $\text{Ln}^{3+}/\text{Ga}^{3+}$ complexes.

Apart from the energy positions of the feeding levels (T_1 , S_1 , CT) with respect to accepting Ln^{3+} levels, another aspect that is usually considered when designing luminescent Ln^{3+} -based coordination compounds is to avoid the presence of O–H, N–H or C–H oscillator in close proximity to Ln^{3+} to prevent the quenching of their excited states. Therefore, the closest distances from Ln^{3+} to these oscillators were estimated for **Ln-1** (Figure S2). It was found that $\text{Ln}\cdots\text{C–H}$ is ~ 7 Å, a value that is close to the one determined for $\text{Ln}^{3+}/\text{Zn}^{2+}$ “encapsulated sandwich” MCs¹¹⁴ and longer than that in $[\text{LnGa}_4(\text{shi})_4(\text{benzoate})_4]^5$ or $[\text{Ln}_2\text{Ga}_8(\text{shi})_8(\text{isophthalate})_4]^3$ (3.4–3.6 Å). However, the presence of O–H groups at ~ 4 Å to Ln^{3+} in the crystal structure of **Ln-1** represents a source of non-radiative deactivations and is probably the reason for the slightly lower Q_{Ln}^{L} values for some of the **Ln-1** in the solid state compared to the $[\text{LnGa}_4(\text{shi})_4(\text{benzoate})_4]^5$ or $[\text{Ln}_2\text{Ga}_8(\text{shi})_8(\text{isophthalate})_4]^3$ MCs. On the other hand, in methanol solutions of **Ln-1**, while the molecular structure is preserved, four OH^- bridges are exchanged with $\text{CH}_3\text{O}^-/\text{CD}_3\text{O}^-$ minimizing the sources of non-radiative deactivations that may explain comparable or higher Q_{Ln}^{L} in respect to those of the aforementioned $\text{Ln}^{3+}/\text{Ga}^{3+}$.

CONCLUSIONS

We have created, synthesized and characterized a new series of $(\text{Ga}^{3+}_4)_2\text{Ln}^{3+}$ MCs (**Ln-1** where $\text{Ln} = \text{Y}^{3+}, \text{Pr}^{3+}, \text{Nd}^{3+}, \text{Sm}^{3+}, \text{Eu}^{3+}, \text{Gd}^{3+}, \text{Tb}^{3+}, \text{Dy}^{3+}, \text{Ho}^{3+}, \text{Er}^{3+}, \text{Tm}^{3+}$ and Er^{3+}) with the general composition $[\text{LnGa}_8(\text{shi})_8(\text{OH})_4]\text{Na}\cdot x\text{CH}_3\text{OH}\cdot y\text{H}_2\text{O}$ that possesses a pseudo- D_{4h} symmetry about the central Ln^{3+} . **Ln-1** in the solid state possess unique photophysical properties. In particular, this is the first example among different series of

Ga³⁺/Ln³⁺ MCs reported so far that exhibits an intense red Eu³⁺-centered emission. Moreover, **Ln-1** contain the first MCs that demonstrate the exceptional ability to sensitize the emissions of six Ln³⁺ in a dual range of wavelengths (visible and NIR: Pr³⁺, Sm³⁺, Dy³⁺, Ho³⁺, Er³⁺ and Tm³⁺). In addition, **Yb-1** and **Nd-1** show characteristic emission in the NIR range while **Tb-1** exhibits bright green luminescence. Therefore, **Ln-1** is one of the few single-molecular complexes able to sensitize f-f transitions of ten out of fifteen trivalent lanthanide(III) ions, excluding La³⁺ (4f⁰), Ce³⁺ (impeding 4f-5d transitions), Pm³⁺ (radioactive), Gd³⁺ (UV emission), and Lu³⁺ (4f¹⁴). In addition, the structure of **Ln-1** remains intact in methanol solutions while four OH⁻ bridges are exchanged with CH₃O⁻/CD₃O⁻ leading to comparable or improved photophysical properties as demonstrated for Ln = Nd³⁺, Sm³⁺, Eu³⁺, Tb³⁺, Dy³⁺, and Yb³⁺. The values of Q_{Ln}^L for **Sm-1** in the visible range are exceptionally high when recorded in both protic (CH₃OH, 10%) and deuterated solvents (CD₃OD, 83%) while those of **Yb-1** are among the highest reported today for Yb³⁺ complexes with non-deuterated and non-halogenated ligands. These results demonstrate that LnGa₈(shi)₈(OH)₄ MCs are able to protect and sensitize Ln³⁺ emission throughout the visible and the NIR ranges opening new and broad perspectives for the exciting applications requiring such exceptional properties. This molecular system offers for the first time a large choice of lanthanide(III) emitters operating simultaneously in dual visible and NIR ranges, providing even more unique spectroscopic signatures.

ASSOCIATED CONTENT

Supporting Information. The Supporting Information is available free of charge at <http://pubs.acs.org>. Experimental details, synthesis and characterization of **Ln-1** MCs, X-ray crystallographic parameters, including CIF file (CCDC 2008321), ESI mass-spectra, NMR studies, supplementary tables and figures describing photophysical properties.

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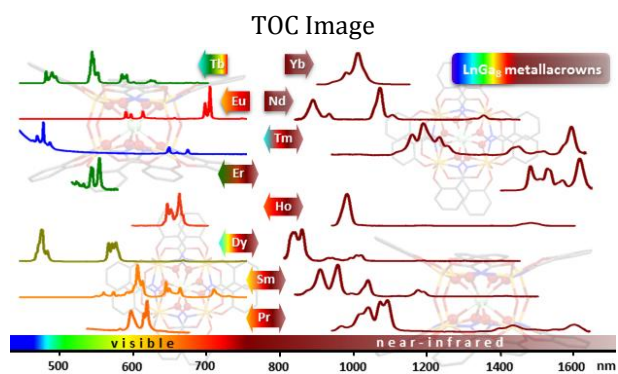
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A sandwich-type metallacrown has been found to sensitize visible, near-infrared, or dual-range emissions from 10 lanthanide(III) ions. Synthesis, structural analysis and extensive photophysical characterization are presented.