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# ***In situ* surface imaging: high temperature Environmental SEM**

## **study of the surface changes during heat treatment of an Al-Si coated boron steel**

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**Abstract:** Al-Si coated boron steels have been of interest for many years because of their superior mechanical properties and their excellent oxidation resistance. These high strength steels are particularly used in hot stamping processes during which the Al-Si coating undergoes multiple microstructural transformations. In this study, morphology and structure transformations are investigated using an original approach. Most of the time, the coating evolution is studied in cross section, by *ex situ* characterization of its different layers after heat treatment. In this work, the evolution of the surface is explored for the first time using *in situ* high-temperature Environmental Scanning Electron Microscopy (HT-ESEM). The austenitization step reproduced in the ESEM chamber allows a precise description of several surface changes occurring, depending on the temperature range. Among them, the main change in morphology is occurring between 650 and 800 °C and is linked to different reactions between aluminium and bi- and ternary Fe-Al-Si phases. The heating rate is also pointed out as a key parameter that affects the surface morphology. In fact, with low heating rates the surface is mostly composed of hexagonal and rectangular elements coming from  $\tau_5$  ( $\text{Fe}_2\text{Al}_8\text{Si}$ ) structuration, while needle-shaped  $\text{FeAl}_3$  structures are found for higher heating rates. In addition, it is observed that the heating rate also affects the surface roughness depending on the surface morphology. Finally, the origin of the presence of micrometric pores appearing between 800 and 900 °C at the coating surface is discussed. Our hypothesis supported by *ex situ* characterization is that they result from formation of a crystallized oxide layer, probably close to  $\alpha\text{-Al}_2\text{O}_3$  structure.

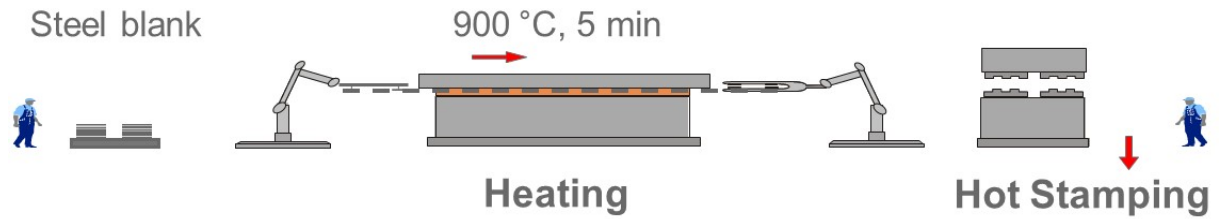
**Keywords:** HT-ESEM; Al-Si boron steel; 3D surface topography; hot stamping; austenitization; intermetallic phases; surface oxidation

## 1. Introduction

Since the last decades, important efforts have been made in the automotive sector to reduce the atmospheric CO<sub>2</sub> emissions and one current approach is to reduce the fuel consumption by lightening the vehicles. In this context, ultra-high strength steels have been developed to minimize their amount, the most common being composed of boron and achieving tensile strengths of around 1500 MPa in quenched condition [1]. These materials, produced using a hot stamping process [2,3], have been significantly employed into vehicles since then, particularly in their body structure, such as door beams, bumpers or roof pillars. In 1995, they accounted for 1 wt.% of the body structure components, while they can now reach up to 60 wt.% [4]. As illustrated in **Figure 1**, in order to achieve these mechanical properties, the steel blank is first heated at 900 - 950 °C for 3 - 10 min in a furnace in order to obtain the austenitic structure. Then, the steel is quenched, leading to a martensitic microstructure [5]. However, under the austenitization conditions, the hot steel surface reacts with air, leading to oxide scale formation and decarburization (oxidation of carbides into CO and/or CO<sub>2</sub>, resulting in a decrease of the carbon content and a loss of the wear resistance) [1].

In order to prevent decarburization as well as oxidation, boron steels are pre-coated with a protective layer. At present, Al-Si are the most widespread coatings. They are obtained by dipping the unprotected steel into an aluminizing bath composed of molten Al, Si and Fe. After aluminizing, the Al-Si coating is composed of two different layers [6]. The thinnest layer (3 μm thick for a usual AS-150 coated steel) is localized at the steel/coating interface and results from the dissolution of iron from the steel into the coating. Its diffusion through the coating yields in an intermetallic Fe-enriched layer composed of Fe-Al binary phases (mostly Fe<sub>2</sub>Al<sub>5</sub> and FeAl<sub>3</sub>, FeAl<sub>3</sub> being also designated as Fe<sub>4</sub>Al<sub>13</sub> in literature [7]), and τ<sub>5</sub> Fe-Al-Si ternary phase. Regarding the τ<sub>5</sub> composition, several propositions have been reported, including Fe<sub>3</sub>Al<sub>12</sub>Si<sub>2</sub> and/or Fe<sub>5</sub>Al<sub>20</sub>Si<sub>2</sub> [8], Fe<sub>3-6</sub>Al<sub>12-15</sub>Si<sub>2-5</sub> [9] or Fe<sub>2</sub>Al<sub>7-8</sub>Si [5,10]. The

second part of the coating, thicker (up to 25  $\mu\text{m}$  thick for a usual AS-150 coated steel), is mainly composed of pure Al, Si rods and Fe-Al-Si ternary phases such as  $\tau_5$  or  $\tau_6$  ( $\text{Fe}_2\text{Al}_9\text{Si}_2$ ).



**Figure 1.** Schematic illustration of a typical hot-stamping line [11].

During austenitization, the Al-Si coating undergoes many transformations, as reported in numerous studies [6,9,10,12–14]. For instance, Wang and Chen [12] studied the oxidation behaviour of Al-Si coated steels between 750 and 950  $^{\circ}\text{C}$  and observed that a high dwell time results in the diffusion of more iron from the steel to the surface coating, accompanied by intermetallic phase transformations. Among these phases, the formation of binary Fe-Al phases such as  $\text{Fe}_4\text{Al}_{13}$ ,  $\text{Fe}_2\text{Al}_5$ ,  $\text{FeAl}_2$  and  $\text{FeAl}$ , was reported [15,16]. Grigorieva *et al.* [6] studied the phase transformations occurring during an austenitization step at 900 - 930  $^{\circ}\text{C}$  by Energy Dispersive X-ray Spectrometry (EDS) analysis of a cross-section, after interrupting the heat treatment at several temperatures. At 615  $^{\circ}\text{C}$ , they observed that the  $\tau_5$  ternary phase, localized at the steel/coating interface, grows rapidly into the coating because of the iron diffusion into molten Si. When the temperature reaches 830  $^{\circ}\text{C}$ , the diffusion of more iron leads to the appearance of more  $\tau_5$  at the coating surface. Moreover, the molten Al can cause  $\text{FeAl}_3$  formation by reaction with  $\tau_5$ . After 10 min of austenitization at 900  $^{\circ}\text{C}$ , they identified five different layers: two layers composed of ternary phases, sandwiched by layers of binary Fe-Al phases. In contrast, Guo *et al.* observed that the coating is composed of only four layers after a heat treatment at 930  $^{\circ}\text{C}$  for 5 min [17]. By EDS analysis, they attributed the surface layer to  $\text{Fe}_2\text{Al}_5$  and  $\text{FeAl}_2$  and the second layer as a mix of ternary phases. In brief, depending on the heating parameters, the coating will be structured differently.

Most of the studies characterizing the Al-Si coating evolution during austenitization are carried out *ex situ* on cross-sections of the coated steel, i.e. only after the heat treatment [3,13,17]. However, to our knowledge, few studies have been focused on the surface evolution during austenitization. Liang *et al.* [14] noted by 3D reconstructions (with optical microscopy) of the coated surface that a maximum surface roughness is obtained at near 700 °C. Moreover, it is known that an alumina layer is formed at the extreme surface by oxidation with air. As mentioned in the review of Fan and De Cooman [5], this oxide layer is often mentioned but without any detailed characterization. Nevertheless, the structure and the composition of the oxide surface are parameters which require in depth analysis for better understanding of surface properties, such as weldability, paintability, corrosion resistance or hydrogen diffusion.

The aim of this work is the identification of the different surface changes at the surface of an Al-Si coating occurring during austenitization. To the best of our knowledge, this is the first time that *in situ* High-Temperature Environmental Scanning Electron Microscopy (HT-ESEM) is used to observe the changes in surface morphology of Al-Si coated steels in real-time during heating. We demonstrate here that such a powerful and fast technique can be used to better understand the parameters affecting the surface thus allowing to better control its properties. The heating conditions are the main parameters investigated in this study, particularly the heating rate. In addition, *ex situ* techniques (XPS, XRD) are used with the aim to characterize the intermetallic phases and the oxide layer at the coating surface.

## 2. Materials and Methods

### 2.1. Material

The material chosen in the present study was a patented 22MnB5 steel covered by a 25  $\mu\text{m}$  thick Al-Si coating [18]. The chemical composition of the boron steel is shown in **Table 1**. The coating was obtained by dipping the steel into an aluminizing bath composed of 88 wt%

Al, 9 wt% Si (Si/Al = 0.1) and 3 wt% Fe at about  $675 \pm 5$  °C. During immersion, part of the iron from the steel dissolves into the coating. After cooling down to ambient temperature,  $Fe_xAl_y$  binary and  $Fe_xAl_ySi_z$  ternary phases form on the coating surface.

**Table 1.** Chemical composition of the 22MnB5 steel

Element	C	Mn	Si	Cr	Ti	B	S	P	Fe
wt.%	0.22	1.3	0.25	0.15	0.04	0.003	0.002	0.013	bal.

In the industrial process, an austenitization step, which consists of heating the coated steel in a furnace at 900 - 930 °C for 3 – 10 min before hot stamping, is performed to obtain the desired mechanical properties. In these real conditions, the starting heating rate is close to 600 °C.min<sup>-1</sup> until a temperature break around 600 °C. From 600 to 900 °C, the heating rate is lower, around 200 °C.min<sup>-1</sup>. During austenitization, further iron from the steel bulk diffuses into the coating, leading to the formation of binary and ternary phases with a higher iron content. In **Table 2**, the most commonly observed phases in Al-Si coatings are listed, accompanied by their crystallographic data.

**Table 2.** Description of the most common solid phases identified in Al-Si coatings on boron steels.

Symbol	Formula	Crystal system, symmetry	Lattice parameters (nm)	JCPDS number	Density	Ref(s)
$\theta$	$FeAl_3$ ( $Fe_4Al_{13}$ )	Monoclinic, C2/m	$a = 1.549, b = 0.808, c = 1.247$ $\beta = 107.7^\circ$	050- 0797	3.8	[7,19]
$\eta$	$Fe_2Al_5$	Orthorhombic, Cmcm	$a = 0.766, b = 0.642, c = 0.422$	029- 0043	4.1	[19– 21]
$\tau_2$	$FeAl_3Si$	Monoclinic  Cubic	$a = 1.780, b = 1.025, c = 0.890$ $\beta = 132^\circ$  $a = 1.603$	020- 0032		[22]
$\tau_5$	$Fe_3Al_{12}Si_2/$ $Fe_5Al_{20}Si_2$ ( $Fe_2Al_{7.4}Si$ )	Hexagonal, P6 <sub>3</sub> /mmc  Cubic	$a = 1.240, c = 2.623$  $a = 1.254$	041- 0894	3.6 – 3.7	[8]
$\tau_6$	$Fe_2Al_9Si_2$	Tetragonal  Monoclinic,	$a = 0.618, c = 4.250$  $a = 0.612, b = 0.612, c = 4.150$	  054-	3.3 – 3.6	[20,23 ]

		A2/a	$\beta = 91^\circ$	0376		
$\omega$	FeSi <sub>2</sub>	Orthorhombic, Cmca	a = 0.987, b = 0.779, c = 0.783	020- 0532		[24]

## 2.2. Characterization

**X-ray Photoelectron Spectroscopy.** X-ray Photoelectron Spectroscopy (XPS) analysis were carried out using an Omicron Argus X-ray photoelectron spectrometer with a monochromated Al K $\alpha$  radiation source ( $h\nu = 1486.6$  eV). The photoelectrons from the sample were analysed at a take-off angle of  $45^\circ$ . The spectra were recorded at room temperature under ultra-high vacuum conditions ( $10^{-9} - 10^{-10}$  mbar), with a 100 eV pass energy for the survey and a 20 eV pass energy for the C 1s, O 1s, N 1s, Al 2p, Si 2p, Zn 2p and Ca 2p regions. Binding energies were calibrated by centring the C 1s peak resulting from adventitious carbon contamination at 284.8 eV and element peak intensities were corrected by Scofield factors [25]. The peak areas were determined after subtraction of a linear background. The spectra were fitted using Casa XPS v.2.3.15 software (Casa Software Ltd., U.K.) and applying a Gaussian/Lorentzian ratio G/L equal to 70/30.

**X-Ray Diffraction.** X-Ray Diffraction (XRD) patterns were recorded using two diffractometers. Austenitized samples were classically analysed with a Bruker D8 Advance powder diffractometer operating at 40 kV and 30 mA in the  $\theta$ - $2\theta$  mode with Bragg-Brentano geometry (few microns of depth analysis), using Cu K $\alpha$  radiation ( $\lambda = 1.5418$  Å). XRD patterns were recorded from  $2\theta=10$  to  $80^\circ$  at a scan rate of  $0.01^\circ \text{ s}^{-1}$ . In order to characterize the thin layers on the surface, in-plane Grazing Incidence X-Ray Diffraction (GIXRD) patterns were collected with a Smartlab Rigaku diffractometer operating at 45 kV and 200 mA, equipped with a 9 kW Cu rotating anode. Data were collected with an incident beam angle of  $0.5^\circ$ , from  $2\theta = 20$  to  $80^\circ$  at a scan rate of  $0.033^\circ \text{ s}^{-1}$ .

**Scanning Electron Microscopy.** Scanning Electron Microscopy (SEM) observations were performed using two different microscopes, one for *ex situ* observation at room-temperature and another one for *in situ* high-temperature measurements and 3D reconstruction.

*Ex situ* room-temperature measurements were performed on a Hitachi SU-70 FEG-SEM operating at 10 kV with a working distance near 15 mm, and equipped with an in-lens secondary electron detector (SE-Lower) for high lateral resolution images. This microscope was particularly used to examine and determine the chemical composition of the surfaces (except the elements lighter than carbon), by the use of an Oxford X-Max detector for localized analysis or elemental mapping.

*In situ* High-Temperature Environmental Scanning Electron Microscopy experiments (HT-ESEM) and 3D reconstructions were conducted using a FEI QUANTA 200 FEG-SEM microscope. For HT-ESEM experiments, the microscope was operating at 20 kV, with a working distance near 22 mm and equipped with a gaseous secondary electron detector (GSED). An *in situ* platinum heating stage allows heating 6 mm diameter pieces of steel with a temperature accuracy of  $\pm 5$  °C, through the use of a homemade thermocouple placed below the sample [26]. Many images (1 per second) of the sample surface were recorded during the heat treatment. The resulting video was then obtained by superimposing the images using ImageJ software [27].

All experiments were performed under 200 Pa of outdoor air. According to the hygrometric data, the experiments were realized with a dew point close to 12 °C. To examine the surface changes occurring during austenitization, Al-Si coated boron steels samples were heated to 900 °C with a dwell time of 5 minutes before cooling down to room temperature (without imposed cooling conditions). In total, four videos of high temperature treatments of the Al-Si coated steels were recorded, with heating rates equals to 20, 50, 100 and 200 °C min<sup>-1</sup> (provided in the Supporting Information). Note that in the industrial conditions the starting



heating rate is close to  $600\text{ }^{\circ}\text{C min}^{-1}$  until a temperature break around  $600\text{ }^{\circ}\text{C}$ . From  $600$  to  $900\text{ }^{\circ}\text{C}$ , the heating rate is lower, around  $200\text{ }^{\circ}\text{C min}^{-1}$ . Since the present work is focusing on what happens from  $600\text{ }^{\circ}\text{C}$  (where the surface transformations occur) then the  $200\text{ }^{\circ}\text{C min}^{-1}$  heating rate fits well with the real industrial conditions. Another experiment was achieved by inserting the raw material into the crucible preheated at  $900\text{ }^{\circ}\text{C}$ . The goal of this experiment was to get even closer to the industrial conditions. Unfortunately, the heating was too fast to record a video, especially due to the time between the closure of the ESEM chamber and the detector ignition (vacuum needed). Only micrographs of the final microstructure were thus recorded and the heating rate was estimated around  $400\text{ }^{\circ}\text{C.min}^{-1}$ . Note also that the quality of the video recorded at heating rates of  $100$  and  $200\text{ }^{\circ}\text{C.min}^{-1}$  is relatively poor, because of the significant mobility of the surface during these experiments and the short frame time used to record the images.

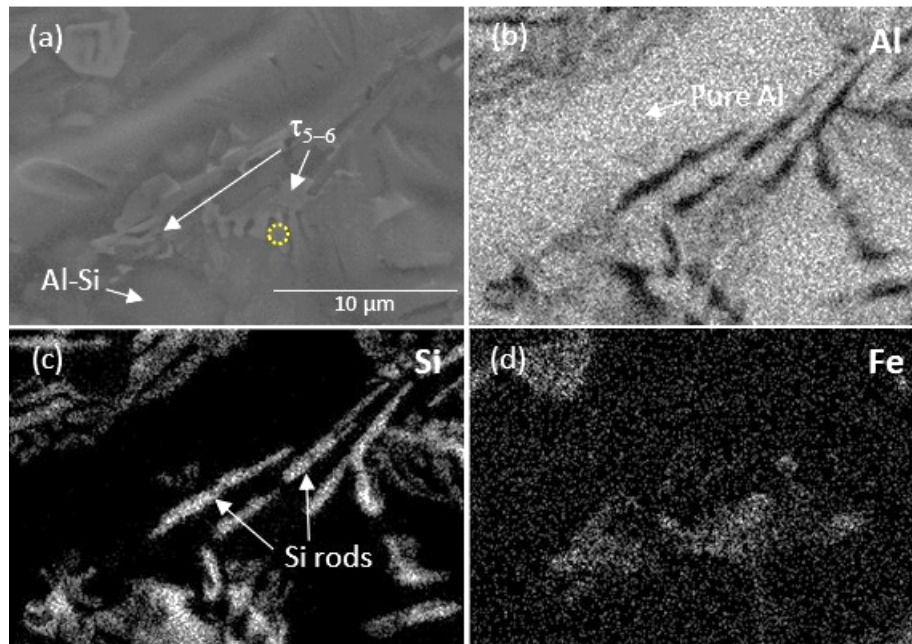
The temperature indicated on videos is not accurate and was corrected by measuring the temperature with a thermocouple directly in contact with the sample. The corrected temperature is provided with each video (**Figures S1 to S4**). Logically, the error percentage between the temperature measured by the thermocouple and the one indicated on the video increases with the heating rate: from 3 to 9% in the  $500 - 900\text{ }^{\circ}\text{C}$  temperature range at  $20$  and  $50\text{ }^{\circ}\text{C.min}^{-1}$ , and from 2 to 25% at  $100 - 200\text{ }^{\circ}\text{C min}^{-1}$ .

This microscope was also used for 3D surface reconstruction (operating at  $2.1\text{ kV}$ ), by recording several images of the same surface with different tilt angles. Once the images were acquired, surface 3D reconstructions were generated using the MeX software (Alicone, Austria) [28]. From the reconstructions, surface roughness parameters were extracted, such as the height average (Sa) and the maximum height (Sz) of the selected area. In addition, the difference in altitude could be locally explored using ImageJ software (version 1.51) [27].

### 3. Results and Discussion

#### 3.1. Characterization of the raw material surface

A micrograph of the raw surface recorded at a x5000 magnification is presented in **Figure 2** (as-received, before austenitization) accompanied by a mapping of the elements Al, Si and Fe. Channels of pure aluminium were identified at the surface, surrounding different phases. Regarding the Si map (**Figure 2c**), several silicon rods were identified but they were not clearly seen on the micrograph, which suggests that they are mostly localized below the extreme surface. At the bottom left of the micrograph an area composed of Al and Si, without Fe is observed. On the other hand, iron can be observed with Al and Si, as intermetallic Fe-Al-Si phases (see arrows on **Figure 2a**). Localized EDS analysis of a dendritic morphology present in the centre of the image (yellow dotted circle on **Figure 2a**), shows that the measured chemical composition corresponds to 13 Al, 2 Fe and 1 Si atoms, thus probably corresponding to a  $\tau_5$  ternary phase ( $\text{Fe}_2\text{Al}_7\text{Si}$ ) plus metallic Al. This result is in agreement with the attribution reported by Maitra and Gupta for dendrites in a Fe-Al-Si ternary system [29]. Furthermore, they also observed that these dendrites were not only composed by the  $\tau_5$  phase but also by the  $\tau_6$  phase ( $\text{Fe}_2\text{Al}_9\text{Si}_2$ ), crystallizing with needle or rod shapes around  $\tau_5$ .



**Figure 2.** Elemental mapping of the raw Al-Si coated 22MnB5 steel surface before austenitization: (a) SEM micrograph at a x5000 magnification and mapping of (b) Al, (c) Si and (d) Fe.

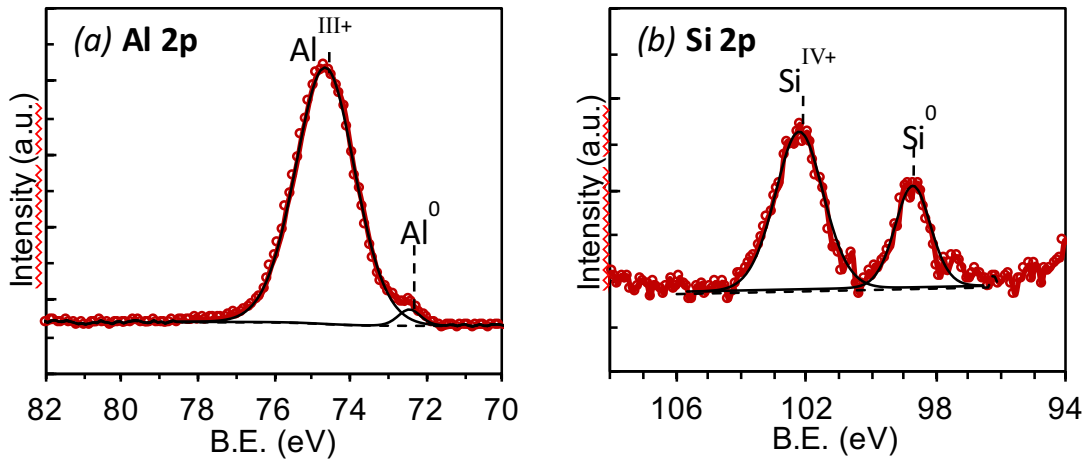
To go further into the description of the raw material, the surface has been characterized by XPS. This technique allows one to obtain an elemental analysis of the first 10 nm of the extreme surface. As presented in **Table 3**, the coating surface was composed of Al, Si, C, O, Ca, Zn and N. Among these elements, carbon and oxygen were the most concentrated. Carbon is mainly of hydrocarbon nature originating from adventitious contamination while oxygen is of hydrocarbon and inorganic nature. The coating was also composed of Al (17.5 at.%) and Si (1.6 at.%). Impurities such as Zn and Ca were also observed and supposed to come from factory pollutions. From the atomic percentages of Al and Si a Si/Al ratio equal to 0.09 was obtained. This result is in agreement with the aluminizing bath composition (Si/Al = 0.1). Interestingly, no trace of Fe was found while  $\tau_5$  ternary phases were observed in **Figure 2**. This difference can be explained by the analysis depth of SEM and XPS which are very different (1000 nm vs. 10 nm respectively). Al 2p and Si 2p peaks, presented in **Figures 3a and 3b**, both showed two contributions due to metallic and oxidized Al and Si compounds,

respectively. These contributions indicate the formation of an oxide layer at the surface of the Al-Si coated 22MnB5 steel.

**Table 3.** Elemental composition of the raw Al-Si coated surface by XPS.

Element	O	C	Si	Al	Zn	N	Ca
At.%	29.2	47.0	1.6	17.5	0.6	1.0	3.1

The absence of a Fe component by XPS analysis can be explained by the presence of an oxide layer above the Fe-Al and Fe-Al-Si layers. Furthermore, by assuming an homogeneous  $\text{Al}_2\text{O}_3$  layer at the sample surface (and neglecting  $\text{SiO}_2$ ), it can be estimated from the Al2p (oxidic) and Al2p (metallic) intensities together with the inelastic mean free paths  $\lambda_{\text{Al}_2\text{O}_3}$  and  $\lambda_{\text{Al}}$  (numerical values from [30]) and the ratio of the volume densities of aluminium in metal to oxide that the alumina layer is around 9 nm thick (thickness calculation detailed in [31]).

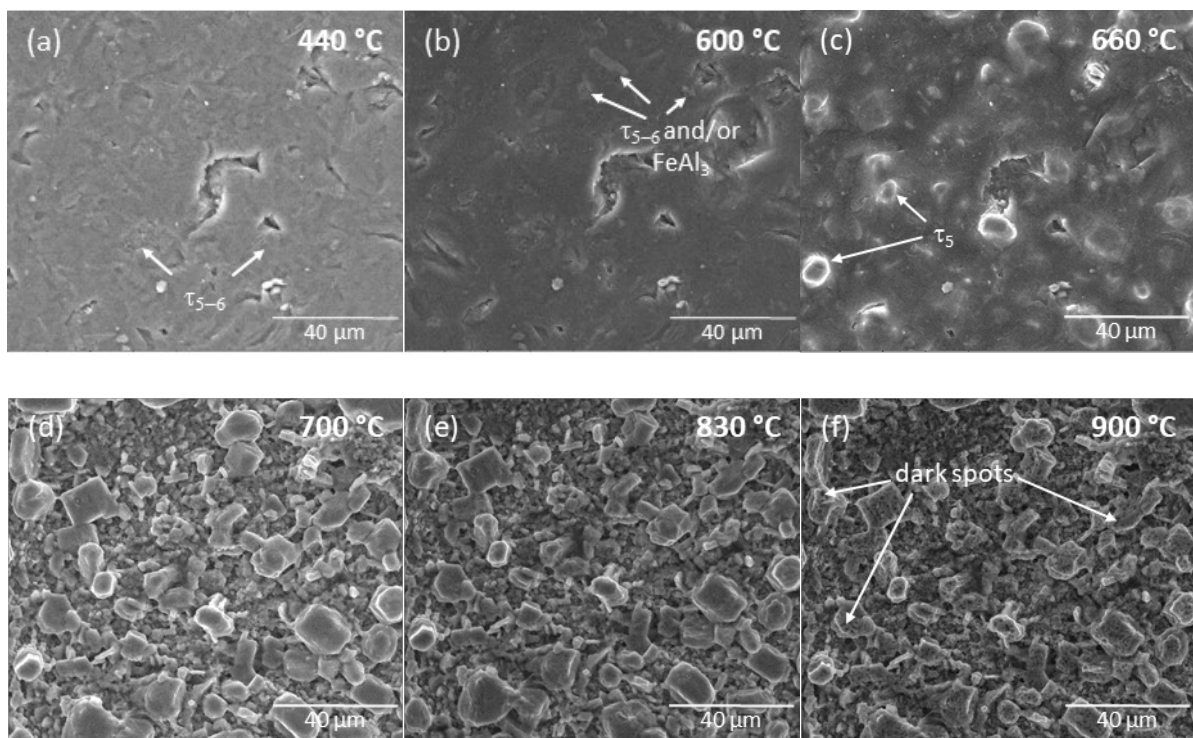


**Figure 3.** XPS Al 2p (a) and Si 2p (b) peaks recorded on the raw Al-Si coated 22MnB5 steel surface (before austenitization).

To summarize, before austenitization, the coating surface is composed of a thin oxide layer of about 10 nm (mainly  $\text{Al}_2\text{O}_3$ ) covering a thicker heterogeneous layer, composed of monometallic Al and Si phases, Al-Si alloys and  $\tau_{5-6}$  ternary phases.

### 3.2. ESEM investigations of the surface changes during heating (austenitization)

ESEM was used to examine the surface changes occurring during a heat treatment simulating the austenitization process. A first heat treatment was realized under ambient air with a low heating rate ( $20\text{ }^{\circ}\text{C min}^{-1}$ ) to identify more easily the different surface changes. From this video (**Video 1**, available in the Supporting Information), several evolutions can be noticed. Selected views from the video describing the main changes of the surface are presented in **Figure 4**.



**Figure 4.** ESEM images recorded during an in situ heat treatment of the Al-Si coated 22MnB5 steel under 200 Pa of air with a temperature ramp of  $20\text{ }^{\circ}\text{C min}^{-1}$ : (a) 440 °C, (b) 600 °C, (c) 660 °C, (d) 700 °C, (e) 830 °C and (f) 900 °C. See text for explanation of the different arrows.

Firstly, no morphological changes are observed from room temperature to 570 °C (**Figure 4a**) with micrographs similar to that shown on **Figure 2** (but with different magnifications), suggesting the presence of  $\tau_5$  and  $\tau_6$  phases. Then, a rapid transformation occurs at 570 °C: the various morphologies assigned to the presence of Fe-Al-Si phases ( $\tau_5$  and  $\tau_6$ ) and Si rods at the surface disappears rapidly, to the benefit of light contrast objects rising at the surface

(**Figure 4b**) that will be attributed to  $\tau_5$ ,  $\tau_6$  and  $\text{FeAl}_3$  phases in the next section. Some areas of the surface also collapse, resulting in the formation of wide cavities (10 - 30  $\mu\text{m}$  wide). By increasing the temperature the entire surface changes, and the bright objects disappear (between 600 and 620  $^\circ\text{C}$ ), suggesting a melting reaction.

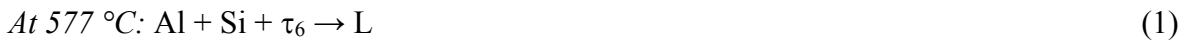
An important surface transformation takes place around 650  $^\circ\text{C}$  characterized by the melting of the surface and the appearance of many structured elements, from hexagonal to rectangular shape (**Figure 4c**) later assigned to  $\tau_5$  phase. This large transformation also reflects the important formation of porosity and consequently the increase of specific surface area. At 700  $^\circ\text{C}$ , this transformation is finally over and the image in **Figure 4d** reveals elements of different shapes and sizes. Among them, the presence of hexagonal and rectangular structures can be underlined with sizes mostly ranging from 5 to 20  $\mu\text{m}$ .

From 780  $^\circ\text{C}$ , the various surface structured elements and particularly the larger ones begin to retract, leading to the appearance of many dark spots on the surface (**Figure 4e**). These spots, smaller than 1  $\mu\text{m}$  in diameter, appears not only on the hexagonal and rectangular structured elements but also on the smallest objects in the background. When the temperature of 900  $^\circ\text{C}$  is reached, dark spots are still appearing (**Figure 4f**) and expands over time, finally covering an important part of the surface at the end of the heat treatment (after 5 min at 900  $^\circ\text{C}$ ) with a final diameter of about 1.5 – 2  $\mu\text{m}$ .

The evolution reported above can be decomposed in four consecutive steps during the heat treatment that can be associated to: (i) dissolution of ternary phases and Si rods starting at 570  $^\circ\text{C}$  (Step 1); (ii) surface liquefaction followed by formation of structured elements with creation of roughness at 650  $^\circ\text{C}$  (Step 2); (iii) retraction of the structured elements at 780  $^\circ\text{C}$  (Step 3) and (iv) formation of dark spots starting also at 780  $^\circ\text{C}$  (Step 4). Although the last two transformations started at the same temperature, they did not end at the same temperature. Indeed, the retraction of the structured elements stopped at 890  $^\circ\text{C}$  while the growth of the

dark spots continued until the end of the heat treatment. The possible explanation is that two different phase transformations occur at the same time. This last step (Step 4) will be further investigated in **part 3.5**.

**Review of phase transformations.** According to a phase transformation study realized by Grigorieva *et al.* [6], three main reactions, selected from the Fe-Al-Si ternary phase diagram established by Raynor and Rivlin [23], are occurring during austenitization of an Al-Si coating. These three reactions ((**1**) to (**3**)) are listed below with their theoretical formation temperature (see **Table 2** for phase description):



Interestingly, one surface transformation observed with HT-ESEM (**Video 1**) is in good agreement with one of the reaction temperatures listed above. According to reaction (**1**), the first transformation recorded at 570 °C in the video should correspond to a reaction involving Al, Si and  $\tau_6$ , leading to their dissolution into a liquid phase. As a matter of fact, the  $\text{Fe}_x\text{Al}_y\text{Si}_z$  phases and Si rods identified on the surface disappear at this temperature. However, in **section 3.1**, the  $\text{Fe}_x\text{Al}_y\text{Si}_z$  objects of dendritic shape were described as a mix of  $\tau_5$  and  $\tau_6$  while, according to reaction (**1**), only  $\tau_6$  reacts with Al and Si at 577 °C. Consequently, the appearance of the bright objects recorded just after the dendrites dissolution (**Figure 4b**) could be attributed to the concentration of residual  $\tau_5$ . This explanation is supported by the fact that some of these bright objects disappear at 620 °C, which corresponds to the temperature of reaction (**2**) involving  $\tau_5$  dissolution in Al matrix. However, some of the objects appearing at 570 °C melt before reaching 620 °C (from 600 °C more exactly). According, to Rivlin and Raynor [23], reaction (**4**) described below, can occur in this temperature range.



We can then conclude that the first step (at 570 °C) cannot be reduced to one reaction but may result from the  $\tau_6$  progressive dissolution into Al, Si or Al-Si with formation of  $\tau_5$  and FeAl<sub>3</sub>. According to reaction (2) at 620 °C, a reaction between the Fe<sub>x</sub>Al<sub>y</sub>Si<sub>z</sub> phase  $\tau_5$  and metallic aluminium causes a transformation into liquidus and the formation of FeAl<sub>3</sub>. At this temperature, bright objects dissolution was recorded on **Video 1**. Even if no new phases that could be attributed to FeAl<sub>3</sub> were distinguished, the latter is still supposed to be present on the surface.

As already described above, the second step leads to the most important surface transformation at about 650 °C consisting in the dissolution of the entire surface with the formation of structured elements. Such transformation is not consistent with any of the three reactions reported by Grigorieva *et al.* [6] or with the other reactions from the Fe-Al-Si ternary phase diagram. However, according to the Fe-Al binary phase diagram [19,32], FeAl<sub>3</sub> dissolution into Al matrix occurs at 655 °C as described by reaction (5), and should explain the important melting of the surface:



FeAl<sub>3</sub> results from reaction (2) at 620 °C by  $\tau_5$  dissolution. Unfortunately, only liquid phase results from the reaction, which cannot explain the formation of structured objects (hexagonal or rectangular). It is interesting to note that among all the quoted phases in **Table 2**, the  $\tau_5$  ternary phase is the only one presenting a hexagonal crystal system. This phase can also crystallize in the cubic system. Thus, it is possible to correlate the hexagonal and rectangular structured elements present on the coating surface (**Figure 4c**) with a  $\tau_5$  ternary phase already formed at 615 °C. Complementary Wavelength Dispersive Spectrometry (WDS) characterizations of the coating surface were conducted after interrupting the heat treatment at different temperatures. Note that cooling step was realized without control of the ramp, which



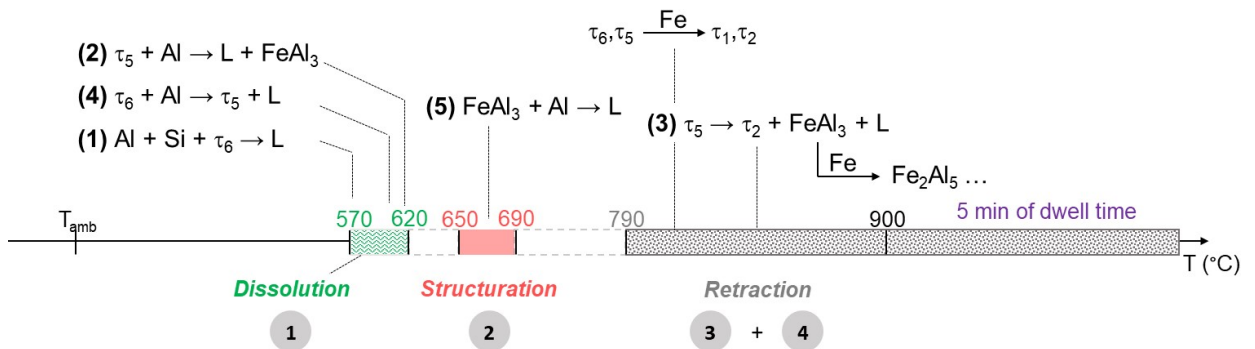
means that RT is rapidly reached. Moreover, no morphological changes were observed during the cooling phase recorded by ESEM. Consequently, even if it could be possible that phase transformations occur locally, it is believed that most of the chemical composition did not change during cooling of the sample. The results conducted after interrupting the heat treatment at 680 °C (see Figure S5) highlighted the presence of  $\tau_5$ ,  $\tau_6$ ,  $\text{FeAl}_3$  and liquidus at the coating surface, thus confirming the reactions (1), (2), (4) and (5) discussed above together with the assumption that no change occurred during cooling.

Regarding the third step of surface transformation (retraction) starting at 780 °C, the temperature difference with reaction (3) is also too important to be correlated only with the last of the main three invariant reactions reported by Grigorieva *et al.* [6]. Complementary WDS characterizations were conducted after interrupting the heat treatment at 810 °C (see Figure S6), namely before the reaction temperature of (3) (855°C). The results especially highlighted the presence of  $\tau_1$  and  $\tau_2$  in addition to  $\tau_5$  and  $\text{FeAl}_3$  on the coating surface. The formation of  $\tau_1$  ( $\text{Al}_3\text{Fe}_3\text{Si}_2$ ) and  $\tau_2$  ( $\text{FeAl}_3\text{Si}$ ) is interesting because of their higher iron content, as compared to  $\tau_5$  and  $\tau_6$ . Thus, their presence could be attributed to a progressive Fe-enrichment of the  $\tau_5$  or  $\tau_6$  phases previously formed by iron dissolution from the steel to the coating surface. Consequently, the beginning of the retraction step could be attributed to their formation before reaction (3) involving  $\tau_5$  dissolution into  $\tau_2$ ,  $\text{FeAl}_3$  and liquidus.

After the retraction step, it is then supposed that the surface is mainly composed of  $\tau_1$  and  $\tau_2$  not reacting in this temperature range after being formed, and  $\text{FeAl}_3$ . Note that according to the binary Fe-Al phase diagram, once  $\text{FeAl}_3$  is formed a local enrichment of iron, caused by its progressive diffusion into the coating at high temperature, can lead to its progressive transformation into  $\text{Fe}_2\text{Al}_5$ ,  $\text{FeAl}_2$  or  $\text{FeAl}$  with an increased Fe/Al ratio, such as previously seen with  $\tau_5$  and  $\tau_6$  transformation into  $\tau_1$  and  $\tau_2$ . Consequently, these last phases can also be present at the surface at the end of the heat treatment.

The principal reason for the retraction of the different structures can be explained by observing the densities of each compound (listed in **Table 2**). For example, the density of the  $\tau_5$  phase is around 3.6 - 3.7 [8] while that of  $\text{FeAl}_3$  is near 3.8 [7]. Moreover,  $\text{Fe}_2\text{Al}_5$  resulting from the Fe enrichment of  $\text{FeAl}_3$  exhibits a higher density, near 4.1 [19,32]. Thus, the transformation of the  $\tau_5$  structures into denser phases could be accompanied by a retraction as observed in ESEM.

The previous discussion can be summarized in the Figure below (**Figure 5**) with all the hypothesized phase transformations all along the four different steps described above that are related to phase transformations. Stabilization periods are also shown, following the dissolution and the structuration steps (Steps 1 and 2 respectively), and correspond to the temperature range during which no transformation is observed at the surface.



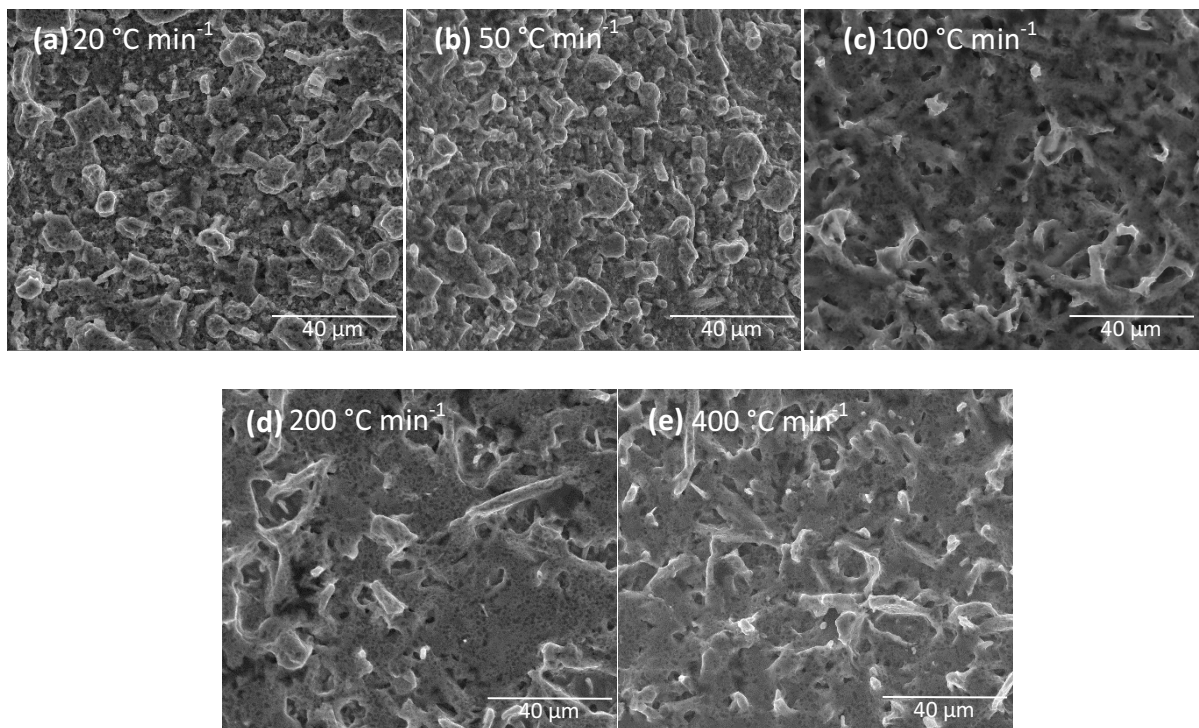
**Figure 5.** Reactions and transformation steps occurring at the surface of an Al-Si coated 22MnB5 steel during a heat treatment under 200 Pa of air up to 900 °C with a temperature ramp of 20 °C min<sup>-1</sup>.

### 3.3. Influence of the temperature rate on the surface morphology

In order to examine the influence of the heating rate on the surface composition and morphology, five different tests were conducted under ambient air up to 900 °C. Heating rates of 20, 50, 100, 200 (**Videos 1 to 4**, Supporting Information) and 400 °C min<sup>-1</sup> were thus compared during HT-ESEM experiments. At first sight, it can be concluded that the same

steps take place during heating. However, some important differences can be seen, such as the final surface morphology or the temperature onsets and temperature ranges of the steps occurring at the surface.

**Figure 6** shows a comparison of the images of each surface obtained after heating at 900 °C, with a temperature ramp between 20 and 400 °C min<sup>-1</sup>. Regarding the initial morphology described in **part 3.2 (Figure 6a, 20 °C min<sup>-1</sup>)**, an important evolution of the morphology was observed by increasing the heating rate. With a heating rate of 50 °C min<sup>-1</sup>, the structured elements at the surface were still hexagonal or rectangular but more rounded (**Figure 6b**) and generally smaller, with a maximum diameter of 15 μm. As detailed in the previous part, the hexagonal and rectangular structures are a consequence of τ<sub>5</sub> formation during heating and consequently, an increased heating rate results in a decline of τ<sub>5</sub> structuration.



**Figure 6.** Influence of the heating rate on the surface morphology after heat treatments at 900°C of the Al-Si coated 22MnB5 steel under 200 Pa of air. Heating rates of (a) 20 °C min<sup>-1</sup>, (b) 50 °C min<sup>-1</sup>, (c) 100 °C min<sup>-1</sup>, (d) 200 °C min<sup>-1</sup> and (e) 400 °C min<sup>-1</sup>.

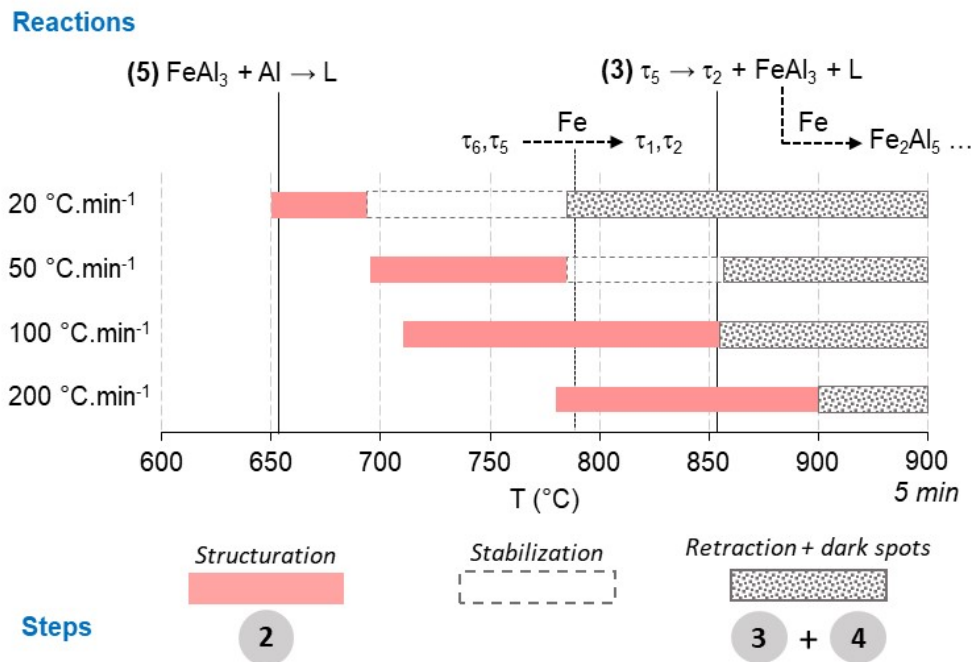
By further increasing the heating rate (**Figure 6c**,  $100\text{ }^{\circ}\text{C min}^{-1}$ ), the surface appears completely transformed after heating up to  $900\text{ }^{\circ}\text{C}$ . The structured elements at the surface are no longer hexagonal or rectangular but rather needle-shaped. The absence of hexagonal and rectangular structuration suggests that  $\tau_5$  ternary phase is not formed in important amounts during structuration in this case. As mentioned by Maitra and Gupta [29], needle-shaped structures can be related to the formation of  $\text{FeAl}_3$  or related compounds. Some wide cavities evidenced by a dark contrast are also present, between the needles. By looking at **Video 3** ( $100\text{ }^{\circ}\text{C min}^{-1}$ , Supporting Information), this important porosity appears between  $820$  and  $850\text{ }^{\circ}\text{C}$  during the heat treatment. When increasing the heating rate ( $200$  and  $400\text{ }^{\circ}\text{C min}^{-1}$ , **Figures 6d and 6e**), the final surface (after completion of the heat treatment) appears still composed of needle-shaped structures, but the needles decrease in number for the benefit of areas without well-defined structures. By considering the contrast of the sample surfaces, it seems that the needles lead to the protrusions of the surface while the poorly defined areas define the valleys. It has to be underlined that the morphology of these final surfaces are very similar to that obtained after a conventional (industrial) austenitization process (see **Figure S7**). Hence, it can be concluded that the two samples shown on **Figures 6d and 6e** are representative of real industrial conditions.

### ***3.4. Correlation between surface morphology and phase transformations***

***In situ HT-ESEM experiments.*** Some temperature ranges related to the four transformation steps (described in **part 3.2**) are shifted to the higher temperatures when increasing the heating rate.

As for the first transformation step, starting with the Al, Si and  $\tau_6$  melting (reaction **(1)**), no difference is observed since the temperatures recorded during the four tests are all included between  $560$  and  $570\text{ }^{\circ}\text{C}$ .

On the contrary, a difference is observed for the second step (structuration of the surface) since this step is shifted to higher temperatures by increasing the heating rate. For example, this transformation step starts around 650 °C when the heating rate is equal to 20 °C min<sup>-1</sup> and shifts to 695 °C and 710 °C with ramps of 50 °C and 100 °C min<sup>-1</sup> respectively. Moreover, compared to the test achieved at 20 °C min<sup>-1</sup>, this step occurs over a longer temperature range for higher heating rates. A summary of the different steps and their shift in temperature from 600 to 900 °C is given in **Figure 7** along with the possible reactions involved (assigned from the experiment realized at 20 °C min<sup>-1</sup>). Note that the dissolution step (Step 1, from 570 to 620 °C) is not detailed in the figure since no difference is observed by increasing the heating rate.



**Figure 7.** Reactions and transformation steps as a function of the heating rate in the 600 – 900 °C range, as defined in **Figure 5**.

From that description, the period of stabilization decreases when increasing the heating rate, as the structuration step is shifted to higher temperatures. With 100 and 200 °C min<sup>-1</sup> heating

rates, no stabilization is observed, the structuration step and the final step occurring successively. With increasing the heating rate, the transformation also shifts to higher temperatures, from 785°C at 20 °C min<sup>-1</sup> to 855 °C at 50 and 100 °C min<sup>-1</sup> and 900 °C at 200 °C min<sup>-1</sup>. All these shifts in temperature cannot be related to a simple kinetic effect since the onset of each step as well as the temperature ranges have no linear relationship with the increase in temperature rate (see for example a similar retraction temperature for 50 and 100 °C min<sup>-1</sup>).

The different reactions possibly occurring in the 600 - 900 °C temperature region shown in **Figure 7** were already described in **part 3.2**. At a heating rate of 50 °C min<sup>-1</sup>, the structuration step is shifted to higher temperature, starting near 700 °C. Thus, reaction (6) (involving FeAl<sub>3</sub> dissolution at 655 °C) is not predominant into the surface transformation. It implies that FeAl<sub>3</sub> was not concentrated at the surface, which could also be explained by less  $\tau_5$  transformation into FeAl<sub>3</sub> in reaction (2) at 620 °C and similarly less  $\tau_6$  transformation into  $\tau_5$  in reaction (4) at 615 °C. In the 600 – 620 °C range, it can be noticed that less bright objects appear at the surface (compare **Video 2** to **Video 1**). Thus, the increase of the temperature rate possibly results in less  $\tau_6$  dissolution which explains the temperature shift of the structuration step. Contrary to the sample heated at 20 °C min<sup>-1</sup>, the retraction step finally started with reaction (3) involving  $\tau_5$  transformation into  $\tau_2$ , FeAl<sub>3</sub> and L.

At 100 °C min<sup>-1</sup>, the structuration step starts at about the same temperature as in the test performed at 50 °C min<sup>-1</sup> but persisted longer. Moreover, **Video 3** (100 °C min<sup>-1</sup>) shows an evolution of surface morphology from  $\tau_5$  hexagonal structures (650 °C) to ill-defined needle-shaped structures when reaching 800 °C. At 855 °C, the formation of long needles was clearly identified resulting from FeAl<sub>3</sub> formation (from reaction (3)). By increasing the temperature, the surface structures keep changing in morphology probably caused by a Fe-enrichment of the present phases. Finally, the retraction step was accompanied by a volume reduction of the

needles on the surface. Unfortunately, the low quality of the video obtained at a high heating rate impedes a precise estimation of the temperature of dark spots appearance. In brief, at  $100\text{ }^{\circ}\text{C min}^{-1}$ , reactions (4) and (5) are inhibited. The surface morphology mainly results of reaction (3), leading to  $\tau_2$  and  $\text{FeAl}_3$  with needle-shaped structuration.

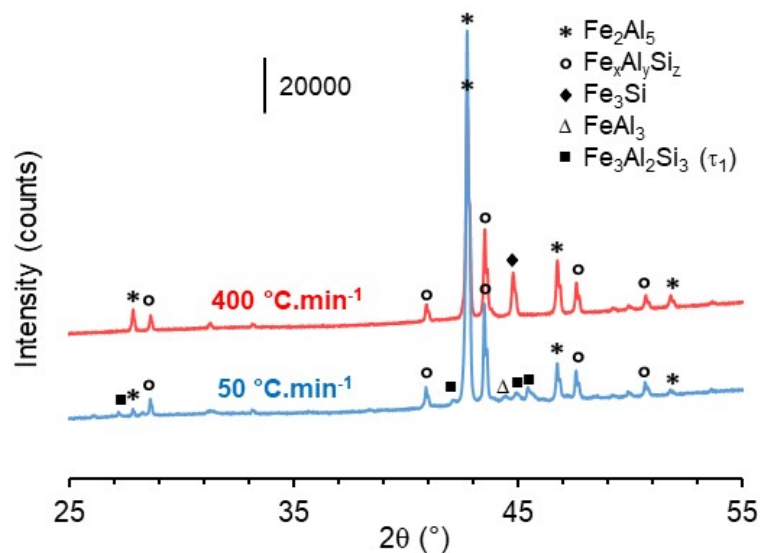
At  $200\text{ }^{\circ}\text{C min}^{-1}$ , the structuration step is again shifted to higher temperatures, in comparison with the three other tests. The surface structuration starts near  $800\text{ }^{\circ}\text{C}$ , probably with  $\tau_5$  and  $\tau_6$  Fe-enrichments leading to  $\tau_1$  and  $\tau_2$  formations. Near  $850\text{ }^{\circ}\text{C}$  in **Video 4**, the appearance of needle-shaped structures ( $\text{FeAl}_3$ , reaction (3)) is clearly visible. Then, the creation of an important porosity starts and continues until  $880\text{ }^{\circ}\text{C}$ . Finally, the retraction step starts after reaching the plateau of  $900\text{ }^{\circ}\text{C}$  probably with a Fe-enrichment of the present phases, especially  $\text{FeAl}_3$ .

To summarize, these results show that the increase of the heating rate causes an increase of the temperature of surface structuration. Depending on this temperature and the associated reactions, the surface morphology changes from  $\tau_5$  hexagonal and rectangular structures to  $\text{FeAl}_3$  needle-shaped structures.

***Ex situ characterizations.*** In order to better understand the surface morphology observed after the heat treatments during the HT-ESEM experiments, the austenitized materials have been characterized using *ex situ* characterization methods after sample cooling to room temperature (5 min of dwell time).

XRD measurements were conducted on materials heated with temperature rates of  $50$  and  $400\text{ }^{\circ}\text{C min}^{-1}$  and the corresponding diffractograms are presented in **Figure 8**. The results show that the two materials do not present exactly the same composition on surface, according to distinct peaks observed in each X-ray diffractogram. Two phases are present on both surfaces:  $\text{Fe}_2\text{Al}_5$  (JCPDS 00-029-0043) and a  $\text{Fe}_x\text{Al}_y\text{Si}_z$  ternary phase. Regarding this last, an attribution

to  $\tau_{12}$  ( $\text{Fe}_2\text{Al}_3\text{Si}$ , JCPDS 04-011-4183) has been reported by Marker *et al.* [10]. They hypothesized that this phase results from a solid-state reaction involving  $\tau_1$  and  $\text{Fe}_2\text{Al}_5$ . Presence of  $\text{Fe}_2\text{Al}_5$  is explained by transformation of  $\text{FeAl}_3$ , mainly resulting from reaction (3) at 855 °C. According to the binary Fe-Al phase diagram, Fe enrichment causes the progressive  $\text{FeAl}_3 \rightarrow \text{Fe}_2\text{Al}_5$  transformation. By increasing the austenitization time, an increasing iron content could also lead to the  $\text{Fe}_2\text{Al}_5$  transformation into Fe richer phases, such as  $\text{FeAl}_2$  and  $\text{FeAl}$  [33].



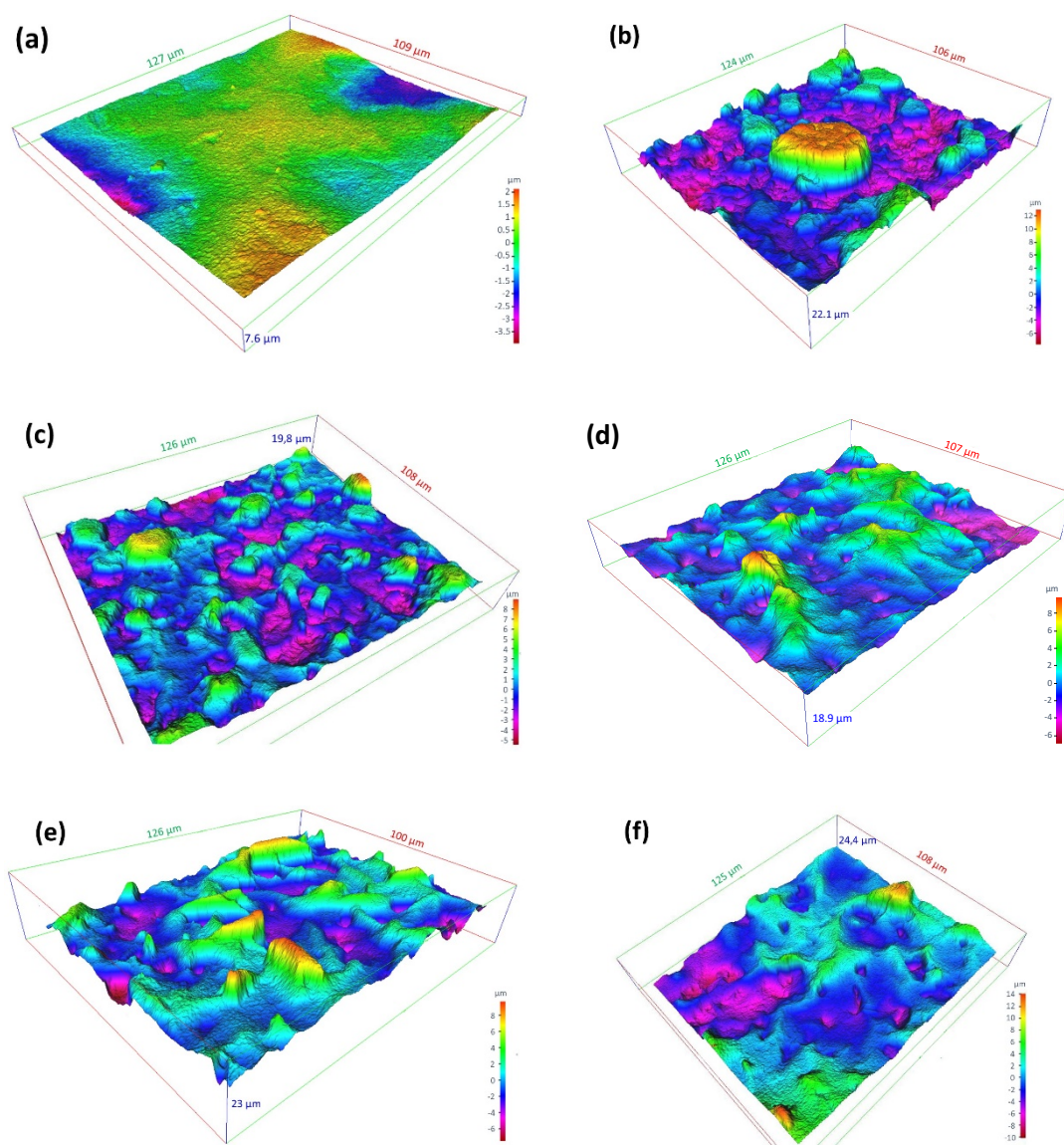
**Figure 8.** XRD profiles of the Al-Si coated 22MnB5 steel, recorded after heat treatments up to 900 °C for 5 min under 200 Pa of air, with a heating rate of (a) 50 °Cmin<sup>-1</sup> and (b) 400 °C min<sup>-1</sup> (tests performed in the ESEM chamber)

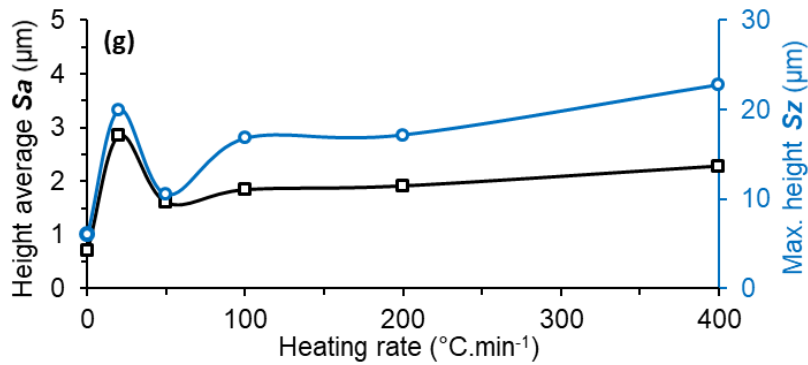
At 50 °C min<sup>-1</sup>, the X-ray diffractogram is also composed of peaks from  $\tau_1$  ternary phase (JCPDS 04-011-4183) and low intense peaks from  $\text{FeAl}_3$  binary phase (JCPDS 00-050-0797), in agreement with the reactions discussed in **part 3.2**. Concerning the treatment performed at 400 °C min<sup>-1</sup>, a  $\text{Fe}_x\text{Si}_y$  phase is also identified, with a peak at 44.8°. This phase has been previously identified by Allely *et al.* [34] as  $\text{Fe}_3\text{Si}$ . This result is to take with caution because



this phase is often observed on the surface of samples having suffered from corrosion during their storage. Hence, its formation may not have occurred during the heat treatment.

3D images of the surface of each sample were reconstructed for a x1000 magnification using both tilted image series and the MeX software [28]. An image of the surface of the raw material before the heat treatment was also reconstructed. The corresponding images are presented in **Figure 9** and accompanied by their surface roughness parameters.





**Figure 9.** 3D reconstructions of the Al-Si coated 22MnB5 steel, prior to (a) and after heat treatments at 900°C (5 min) with heating rates of (b) 20, (c) 50, (d) 100, (e) 200 and (f) 400 °C min<sup>-1</sup> (3D reconstructions using tilted images recorded for a x1000 magnification), and (g) evolution of the corresponding surface roughness parameters (the height average  $S_a$  and the maximum height  $S_z$  in  $\mu\text{m}$ ).

The raw material shows a moderate surface roughness (**Figure 9a**) with a height average ( $S_a$ ) of about 0.7  $\mu\text{m}$  and a maximum height ( $S_z$ ) of 6.1  $\mu\text{m}$  between the peaks and valleys. The heat treatment (**Figures 9b to f**) caused an important increase of the height difference between peaks and valleys of the surface ( $S_z$ ). At 20 °C min<sup>-1</sup> (**Figure 9b**), the maximum height was close to 20  $\mu\text{m}$ , which is more than three times higher than for the raw material. Surprisingly, at 50 °C min<sup>-1</sup> (**Figure 9c**), the surface roughness parameters drop,  $S_a$  being for example two times lower than for 20 °C min<sup>-1</sup>. Then, with heating rates ranging from 100 to 400 °C min<sup>-1</sup> (**Figures 9d to f**), both parameters ( $S_a$  and  $S_z$ ) gradually increased again. At 400 °C min<sup>-1</sup>, the final surface roughness was close to that at 20 °C min<sup>-1</sup> with a  $S_z$  close to 22  $\mu\text{m}$ .

The fact that no clear trend was identified when increasing the heating rate is in contradiction with the study of Liang *et al.* [14], who observed that the surface was gradually rougher when increasing the heating parameters. They assigned this increase of roughness to stronger surface tensions by increasing the temperature rate, caused by less diffusion time and inhomogeneous distribution of intermetallic phases. From 100 to 400 °C min<sup>-1</sup>, without change of surface morphology (needle-shaped structures), the increase of roughness is also observed. However, at lower heating rates (from 20 to 100 °C min<sup>-1</sup>), as discussed above, strong changes in surface morphology were identified, starting with the  $\tau_5$  structuration at 20

$^{\circ}\text{C min}^{-1}$ . At  $50\text{ }^{\circ}\text{C min}^{-1}$ , hexagonal and rectangular structures coming from  $\tau_5$  formation were less structured than at  $20\text{ }^{\circ}\text{C min}^{-1}$ . The 3D reconstruction (compare **Figure 9b and 9c**) also underlines the genesis of needle-shaped structures at  $50\text{ }^{\circ}\text{C min}^{-1}$  between the valleys of the surface. Then, by increasing the heating rate ( $100\text{ }^{\circ}\text{C min}^{-1}$ ) this needle-shaped structuration is more expressed, accompanied with an increase of the surface roughness. Hence, the transition between the  $\tau_5$  ( $20\text{ }^{\circ}\text{C min}^{-1}$ ) and  $\text{FeAl}_3$  ( $100\text{ }^{\circ}\text{C min}^{-1}$ ) structurations causes a decrease in surface roughness followed by an increase. In brief, these results show that the heating rate is not the only parameter to take into account to predict the roughness of an Al-Si surface after austenitization. The surface morphology is also a determining parameter which can explain the non-linear evolution of surface roughness from 20 to  $400\text{ }^{\circ}\text{C min}^{-1}$ .

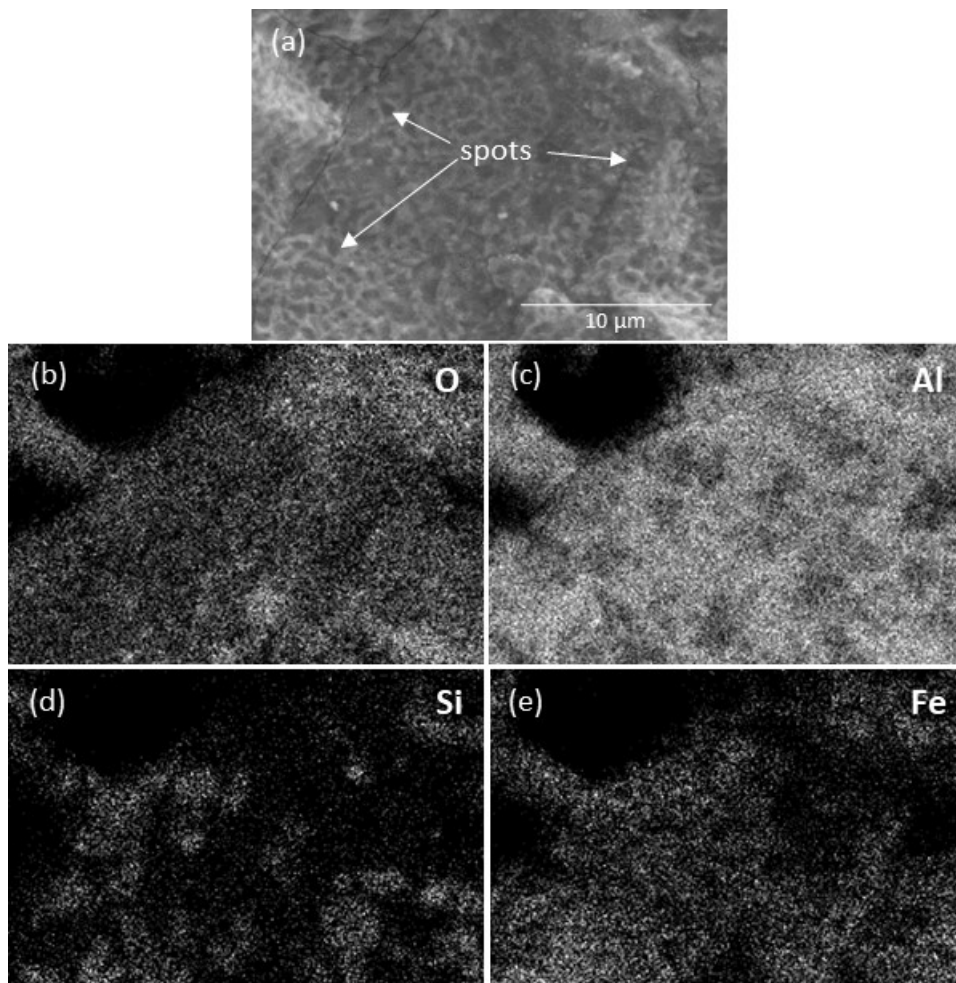
### ***3.5. Identification of dark contrast spots***

In **part 3.2** the appearance of dark spots at the surface has been noticed at the beginning of the retraction step and these spots kept growing in width and densify at the surface until the end of the heat treatment. These spots were found at the surface of each treated sample, whatever the heating rate. In order to determine the origin, composition and structuration of the dark spots, different characterizations were conducted.

Firstly, after a heat treatment at  $200\text{ }^{\circ}\text{C min}^{-1}$  up to  $900^{\circ}\text{C}$ , XPS analysis showed only Al 2p, O 1s and C 1s peaks, while Si 2p was identified on the raw material (see **part 3.1**). Interestingly, the XPS spectrum of the Al 2p photoelectron line (available in Supporting Information, **Figure S8**) shows only one contribution at 74.8 eV, assigned to oxidized aluminium which means that the surface is covered by an alumina layer with a thickness exceeding the analysis depth of XPS (i.e.  $\sim 10\text{ nm}$ ).

Second, elemental mapping of the surface was conducted with EDS coupled to SEM. **Figure 10** shows the image of the surface obtained after a heat treatment at  $200\text{ }^{\circ}\text{C min}^{-1}$  up to  $900^{\circ}\text{C}$  along with the elemental maps of Al, Si, O and Fe. Al, O and Fe are homogeneously

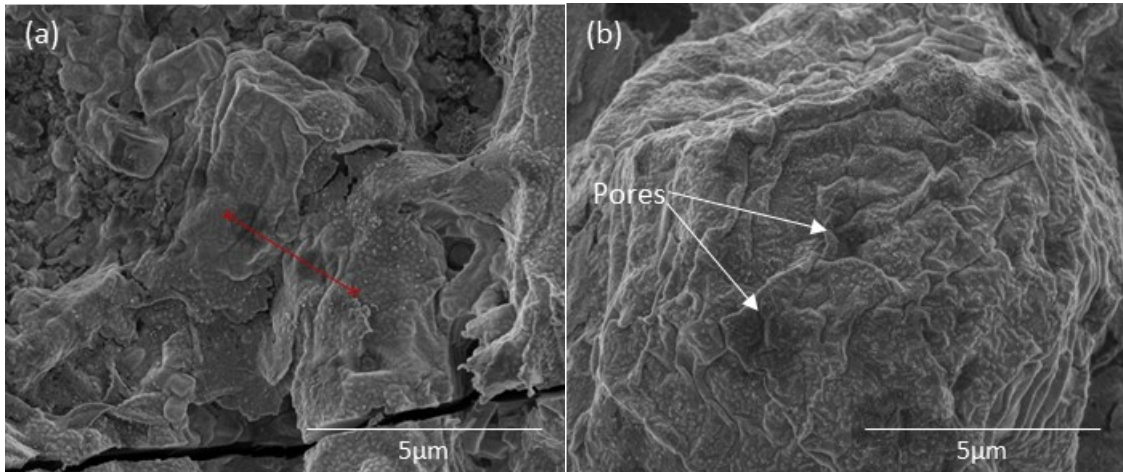
dispersed on the surface while Si is concentrated in areas ranging from 2 to 5  $\mu\text{m}$  wide. Presence of O, in line with the XPS analysis, confirms the formation of an oxide layer. EDS analysis yields an atomic percentage of O between 30 and 40 % and 40-55% for Al. Consequently, it can be concluded that the surface is covered by a thick alumina layer with a thickness of a few hundred nanometers, according to the depth of EDS analysis. By comparing the location of the dark spots with the elemental maps, no direct connection can be made with the presence or absence of one element.



**Figure 10.** Elemental mapping of the Al-Si coated 22MnB5 steel surface after austenitization at 900  $^{\circ}\text{C}$  (heating rate of 200  $^{\circ}\text{C min}^{-1}$ ): (a) SEM micrograph at a x5000 magnification and mapping of (b) O, (c) Al, (d) Si and (e) Fe.

In parallel, tilted images were recorded at high magnification (x25000) in order to permit a careful observation of the surface topography and the local details such as the dark spots

(**Figures 11a and 11b**,  $50\text{ }^{\circ}\text{C min}^{-1}$ ). This processing showed that the two dark spots crossed by the red segment in **Figure 11a** are cavities of  $1.1 - 1.2\text{ }\mu\text{m}$  diameter and  $130 - 160\text{ nm}$  depth. The micrometric pores pointed by an arrow in **Figure 11b** are of the same dimension ( $0.8\text{ }\mu\text{m}$  wide and  $180\text{ nm}$  deep).

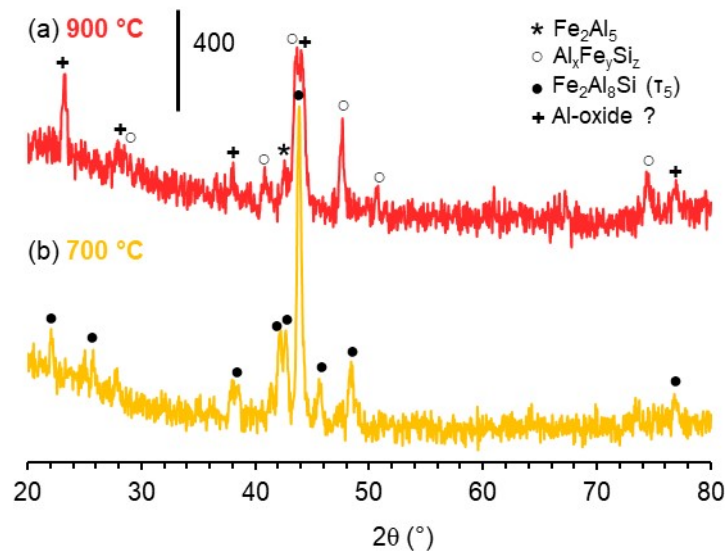


**Figure 11.** Tilted SEM images (a) and (b) ( $\times 25000$  magnification, from two different samples) of the Al-Si coated 22MnB5 steel surface recorded after austenitization at  $900\text{ }^{\circ}\text{C}$  (heating rate of  $50\text{ }^{\circ}\text{C min}^{-1}$ ).

EDS analyses showed that a thick aluminium oxide layer is formed at the sample surface while the pores are around  $130 - 180\text{ nm}$  deep, which means that they are probably located in the alumina layer. Secondary-Ion Mass Spectrometry (SIMS) analysis (results not shown) confirmed that the oxide layer is between  $150$  and  $200\text{ nm}$  thick.

In order to describe the structuration of the oxide layer that should be highly dependent on the environment and heating conditions [35–37], in-plane GI-XRD measurements were recorded to characterize the outermost part of the surface (less than  $1\text{ }\mu\text{m}$ ). Two X-ray diffractograms are shown on **Figure 12** for samples obtained after a classical treatment at  $900\text{ }^{\circ}\text{C}$  (heating rate of  $200\text{ }^{\circ}\text{C min}^{-1}$ , dwell time of  $5\text{ min}$ ) and after inserting the raw coated steel into a muffle furnace preheated at  $700\text{ }^{\circ}\text{C}$  for  $7\text{ min}$ . The goal is to compare the surface crystallinity

of the extreme surface before and after the appearance of dark cavities. Regarding the sample heated at 700 °C, peaks at 22.2°, 42.8°, 43.9°, 45.7° and 48.4° are detected indicating the presence of a  $\tau_5$  ternary phase (JCPDS 00-041-0894), in full agreement with reactions (4) in the 600 – 700 °C temperature range. No peaks ascribed to  $\text{Al}_2\text{O}_3$  or  $\text{SiO}_2$  phases are detected, while the presence of an alumina layer with a thickness larger than 15 nm is confirmed by XPS analysis (results not shown). Two explanations are possible: either the amount of  $\text{Al}_2\text{O}_3$  and  $\text{SiO}_2$  phases is too low to be detected or both phases are amorphous.



**Figure 12.** Grazing Incidence-XRD profiles of the Al-Si coated 22MnB5 steel, recorded after a heat treatment under ambient air at: (a) 900 °C for 5 min, with a heating rate of 200 °C  $\text{min}^{-1}$ ; (b) at 700 °C by inserting the raw material into a muffle furnace preheated (total insertion time of 7 min).

Grazing-XRD analysis of the sample obtained after a heat treatment at 900 °C indicates different phases compared to the previous sample. A peak at 42.7° can be assigned to  $\text{Fe}_2\text{Al}_5$  (JCPDS 00-029-0043), while the peaks at 28.8, 40.8, 43.6, 47.7, 50.7 and 74.5° are ascribed to the  $\text{Fe}_x\text{Al}_y\text{Si}_z$  ternary phase already identified in **Figure 8**. XRD lines at 23.3, 28.0, 38.0, 44.1, and 76.9° cannot be fully assigned to a specific structure. Allély *et al.*[38] characterized the same coated materials by XRD and also identified a peak at 44.1° that was attributed to Al



oxide and more precisely to  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> (JCPDS 00-010-0173). However, as mentioned by the authors, other characteristic peaks of  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> were absent. In our case, peaks at 38.0 and 76.9° are very close to those expected for  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> (37.8 and 76.9°) while peaks at 23.3 and 28.0° are shifted (25.6 and 35.2° for  $\alpha$ -Al<sub>2</sub>O<sub>3</sub>). The fact that peaks are missing or shifted can have several reasons: either the phase is not pure, or the lattice and the crystal structure are altered by mechanical stress. Consequently, according to these results, the formation of a crystallized surface layer close to  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> is strongly assumed but Grazing-XRD cannot bring a definitive answer. However, this hypothesis is also supported by the work of Kitajima *et al.* [36] who demonstrated that  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> layer can easily be formed on a Fe-Al alloy at 900 °C.

Classical XRD measurements on the same sample (**Figure 8**) did not permit identification of  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> which means that this phase is localised at the extreme surface (less than 1  $\mu$ m deep) and in a low amount. Moreover, this phase was only identified when the sample was heated to 900 °C. Consequently, a link between the formation of a surface  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> phase and the genesis of dark spots can be postulated.

To summarize, the dark spots appearing at the beginning of the retraction step (between 780 and 900 °C depending on temperature ramps) are micrometric pores (130 – 180 nm deep) that are most probably a consequence of the progressive formation of crystallized alumina (close to  $\alpha$ -Al<sub>2</sub>O<sub>3</sub>) at the surface. This mechanism explains why these dark spots keep growing with time, even if no phase transformation is occurring.

#### **4. Conclusion**

High-temperature Environmental SEM (HT-ESEM) investigation of the surface transformation of an Al-Si coated 22MnB5 steel up to 900°C in ambient air shows that there

are three important temperature ranges linked to surface changes during austenitization. Near 570 °C, a first evolution is related to the dissolution of Al, Si and  $\tau_6$  phases into liquidus. A second transformation takes place between 650 and 800 °C (according to the temperature ramp) and is the most important in terms of final surface morphology. This evolution can be rationalized with several reactions between ternary phases, binary  $\text{Fe}_x\text{Al}_y$  phases and liquid Al. Moreover, the type of reaction occurring is depending on the heating rate, which has a drastic effect on the surface structuration and roughness. At low heating rates ( $\leq 50 \text{ }^\circ\text{C min}^{-1}$ ), the surface at 900°C was mostly composed of hexagonal and rectangular elements coming from  $\tau_5$  ( $\text{Fe}_2\text{Al}_8\text{Si}$ ) structuration. At high heating rates ( $\geq 50 \text{ }^\circ\text{C min}^{-1}$ ) the surface at 900°C is made of needle-shaped  $\text{FeAl}_3$  structures. Finally, the third and last surface transformation takes place between 800 and 900 °C and is mainly explained by the transformation of the previous surface phases into denser phases, resulting in a retraction of the structures present at the surface. In parallel, the identification of dark spots appearing at the end of the heating step at 900°C, could be attributed to the formation of crystallized  $\alpha\text{-Al}_2\text{O}_3$  leading to the genesis of cavities and opening a porosity in the surface of the coated steel. The use of HT-ESEM proves to be a powerful *in situ* tool to understand the surface transformations occurring on the surface of Al-Si coatings. It has the advantage to yield a direct and fast correlation between the temperature and a specific transformation, which is quite difficult to achieve using *ex situ* characterization. Moreover, it allows one to identify rapid transformations such as those occurring between 570 and 600 °C that would otherwise have been difficult to observe.

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### **Data availability**

The raw data required to reproduce these findings (video) are available from the supporting information section.

## References

- [1] H. Karbasian, A.E. Tekkaya, A review on hot stamping, *J. Mater. Process. Technol.* 210 (2010) 2103–2118. <https://doi.org/10.1016/j.jmatprotec.2010.07.019>.
- [2] K. Mori, P.F. Bariani, B.-A. Behrens, A. Brosius, S. Bruschi, T. Maeno, M. Merklein, J. Yanagimoto, Hot stamping of ultra-high strength steel parts, *CIRP Ann.* 66 (2017) 755–777. <https://doi.org/10.1016/j.cirp.2017.05.007>.
- [3] J.N. Rasera, K.J. Daun, C.J. Shi, M. D’Souza, Direct contact heating for hot forming die quenching, *Appl. Therm. Eng.* 98 (2016) 1165–1173. <https://doi.org/10.1016/j.applthermaleng.2015.12.142>.
- [4] T. Taylor, A. Clough, Critical review of automotive hot-stamped sheet steel from an industrial perspective, *Mater. Sci. Technol.* 34 (2018) 809–861. <https://doi.org/10.1080/02670836.2018.1425239>.
- [5] D.W. Fan, B.C. De Cooman, State-of-the-Knowledge on Coating Systems for Hot Stamped Parts, *Steel Res. Int.* 83 (2012) 412–433. <https://doi.org/10.1002/srin.201100292>.
- [6] R. Grigorieva, P. Drillet, J.M. Maigne, A. Redjaïmia, Phase Transformations in the Al-Si Coating during the Austenitization Step, *Solid State Phenom.* 172–174 (2011) 784–790. <https://doi.org/10.4028/www.scientific.net/SSP.172-174.784>.
- [7] P.J. Black, The structure of  $\text{FeAl}_3$ . I, *Acta Crystallogr.* 8 (1955) 43–48. <https://doi.org/10.1107/S0365110X5500011X>.
- [8] R.N. Corby, P.J. Black, The structure of  $\alpha$ -(AlFeSi) by anomalous-dispersion methods, *Acta Crystallogr. Sect. B.* 33 (1977) 3468–3475. <https://doi.org/10.1107/S0567740877011224>.
- [9] C. Georges, T. Sturel, P. Drillet, J.-M. Maigne, Absorption/Desorption of Diffusible Hydrogen in Aluminized Boron Steel, *ISIJ Int.* 53 (2013) 1295–1304. <https://doi.org/10.2355/isijinternational.53.1295>.
- [10] M.C.J. Marker, B. Skolyszewska-Kühberger, H.S. Effenberger, C. Schmetterer, K.W. Richter, Phase equilibria and structural investigations in the system Al-Fe-Si, *Intermetallics.* 19 (2011) 1919–1929. <https://doi.org/10.1016/j.intermet.2011.05.003>.
- [11] Daimler Chrysler, Typical hot-stamping lines, *Automotive Symposium Cannes 2005*.
- [12] C.-J. Wang, S.-M. Chen, The high-temperature oxidation behavior of hot-dipping Al-Si coating on low carbon steel, *Surf. Coat. Technol.* 200 (2006) 6601–6605. <https://doi.org/10.1016/j.surfcoat.2005.11.031>.
- [13] L. Pelcastre, J. Hardell, A. Rolland, B. Prakash, Influence of microstructural evolution of Al-Si coated UHSS on its tribological behaviour against tool steel at elevated temperatures, *J. Mater. Process. Technol.* 228 (2016) 117–124. <https://doi.org/10.1016/j.jmatprotec.2015.03.009>.
- [14] W. Liang, W. Tao, B. Zhu, Y. Zhang, Influence of heating parameters on properties of the Al-Si coating applied to hot stamping, *Sci. China Technol. Sci.* 60 (2017) 1088–1102. <https://doi.org/10.1007/s11431-016-0231-y>.
- [15] M. Windmann, A. Röttger, I. Hahn, W. Theisen, Mechanical properties of AlXFeY intermetallics in Al-base coatings on steel 22MnB5 and resulting wear mechanisms at press-hardening tool steel surfaces, *Surf. Coat. Technol.* 321 (2017) 321–327. <https://doi.org/10.1016/j.surfcoat.2017.04.075>.
- [16] M. Windmann, A. Röttger, W. Theisen, Phase formation at the interface between a boron alloyed steel substrate and an Al-rich coating, *Surf. Coat. Technol.* 226 (2013) 130–139. <https://doi.org/10.1016/j.surfcoat.2013.03.045>.
- [17] M. Guo, K. Gao, W. Wang, X. Wei, Microstructural evolution of Al-Si coating and its influence on high temperature tribological behavior of ultra-high strength steel against H13 steel, *J. Iron Steel Res. Int.* 24 (2017) 1048–1058. [https://doi.org/10.1016/S1006-706X\(17\)30152-8](https://doi.org/10.1016/S1006-706X(17)30152-8).

- [18] J.-P. Laurent, J.-P. Hennechart, D. Spehner, J. Devroc, Coated hot- and cold-rolled steel sheet comprising a very high resistance after thermal treatment, US6296805B1, 2001. <https://patents.google.com/patent/US6296805B1/en?q=FR2780984> (accessed July 18, 2018).
- [19] B. Sundman, I. Ohnuma, N. Dupin, U.R. Kattner, S.G. Fries, An assessment of the entire Al-Fe system including D03 ordering, *Acta Mater.* 57 (2009) 2896–2908. <https://doi.org/10.1016/j.actamat.2009.02.046>.
- [20] V.I. Dybkov, V.P. Dybkov, *Growth Kinetics of Chemical Compound Layers*, Cambridge Int Science Publishing, 1998.
- [21] U. Burkhardt, Y. Grin, M. Ellner, K. Peters, Structure refinement of the iron–aluminium phase with the approximate composition Fe<sub>2</sub>Al<sub>5</sub>, *Acta Crystallogr. Sect. B.* 50 (1994) 313–316. <https://doi.org/10.1107/S0108768193013989>.
- [22] N. Krendelsberger, F. Weitzer, J.C. Schuster, On the Reaction Scheme and Liquidus Surface in the Ternary System Al-Fe-Si, *Metall. Mater. Trans. A.* 38 (2007) 1681–1691. <https://doi.org/10.1007/s11661-007-9182-x>.
- [23] V.G. Rivlin, G.V. Raynor, 4: Critical evaluation of constitution of aluminium-iron-silicon system, *Int. Met. Rev.* 26 (1981) 133–152. <https://doi.org/10.1179/imtr.1981.26.1.133>.
- [24] G. Shao, Y. Gao, X.H. Xia, M. Milosavljević, Crystallographic characteristics and fine structures of semiconducting transition metal silicides, *Thin Solid Films.* 519 (2011) 8446–8450. <https://doi.org/10.1016/j.tsf.2011.05.036>.
- [25] J.H. Scofield, Hartree-Slater subshell photoionization cross-sections at 1254 and 1487 eV, *J. Electron Spectrosc. Relat. Phenom.* 8 (1976) 129–137. [https://doi.org/10.1016/0368-2048\(76\)80015-1](https://doi.org/10.1016/0368-2048(76)80015-1).
- [26] R. Podor, D. Pailhon, J. Ravoux, H.-P. Brau, Development of an Integrated Thermocouple for the Accurate Sample Temperature Measurement During High Temperature Environmental Scanning Electron Microscopy (HT-ESEM) Experiments, *Microsc. Microanal.* 21 (2015) 307–312. <https://doi.org/10.1017/S1431927615000252>.
- [27] C.A. Schneider, W.S. Rasband, K.W. Eliceiri, NIH Image to ImageJ: 25 years of Image Analysis, *Nat. Methods.* 9 (2012) 671–675.
- [28] M. Pirisinu, V. Mazzarello, 3D profilometric characterization of the aged skin surface using a skin replica and alicona Mex software, *Scanning.* 38 (n.d.) 213–220. <https://doi.org/10.1002/sca.21256>.
- [29] T. Maitra, S.P. Gupta, Intermetallic compound formation in Fe–Al–Si ternary system: Part II, *Mater. Charact.* 49 (2002) 293–311. [https://doi.org/10.1016/S1044-5803\(03\)00005-6](https://doi.org/10.1016/S1044-5803(03)00005-6).
- [30] S. Tanuma, C.J. Powell, D.R. Penn, Calculations of electron inelastic mean free paths for 31 materials, *Surf. Interface Anal.* 11 (1988) 577–589. <https://doi.org/10.1002/sia.740111107>.
- [31] B.R. Strohmeier, An ESCA method for determining the oxide thickness on aluminum alloys, *Surf. Interface Anal.* 15 (1990) 51–56. <https://doi.org/10.1002/sia.740150109>.
- [32] X. Li, A. Scherf, M. Heilmaier, F. Stein, The Al-Rich Part of the Fe-Al Phase Diagram, *J. Phase Equilibria Diffus.* 37 (2016) 162–173. <https://doi.org/10.1007/s11669-015-0446-7>.
- [33] M. Windmann, A. Röttger, W. Theisen, Formation of intermetallic phases in Al-coated hot-stamped 22MnB5 sheets in terms of coating thickness and Si content, *Surf. Coat. Technol.* 246 (2014) 17–25. <https://doi.org/10.1016/j.surfcoat.2014.02.056>.
- [34] N.R. Baldwin, D.G. Ivey, Iron silicide formation in bulk iron-silicon diffusion couples, *J. Phase Equilibria.* 16 (1995) 300–307. <https://doi.org/10.1007/BF02645285>.
- [35] A. Boumaza, L. Favaro, J. Lédion, G. Sattonnay, J.B. Brubach, P. Berthet, A.M. Huntz, P. Roy, R. Tétot, Transition alumina phases induced by heat treatment of boehmite: An

- X-ray diffraction and infrared spectroscopy study, *J. Solid State Chem.* 182 (2009) 1171–1176. <https://doi.org/10.1016/j.jssc.2009.02.006>.
- [36] Y. Kitajima, S. Hayashi, T. Nishimoto, T. Narita, S. Ukai, Rapid Formation of  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> Scale on an Fe–Al Alloy by Pure-Metal Coatings at 900 °C, *Oxid. Met.* 73 (2010) 375–388. <https://doi.org/10.1007/s11085-009-9184-8>.
- [37] S.R.J. Saunders, M. Monteiro, F. Rizzo, The oxidation behaviour of metals and alloys at high temperatures in atmospheres containing water vapour: A review, *Prog. Mater. Sci.* 53 (2008) 775–837. <https://doi.org/10.1016/j.pmatsci.2007.11.001>.
- [38] C. Allély, L. Dosdat, O. Clauzeau, K. Ogle, P. Volovitch, Anticorrosion mechanisms of aluminized steel for hot stamping, *Surf. Coat. Technol.* 238 (2014) 188–196. <https://doi.org/10.1016/j.surfcoat.2013.10.072>.