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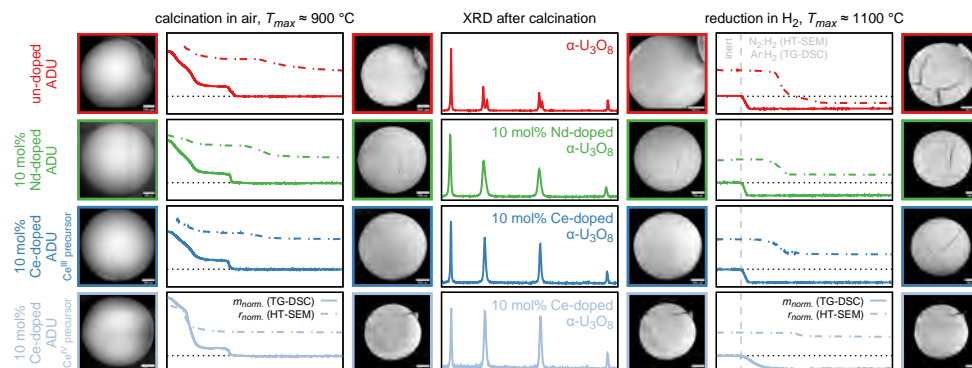
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Graphical Abstract

Structural changes of Nd- and Ce-doped ammonium diuranate microspheres during the conversion to $U_{1-y}Ln_yO_{2\pm x}$

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Highlights

Structural changes of Nd- and Ce-doped ammonium diuranate microspheres during the conversion to $U_{1-y}Ln_yO_{2\pm x}$

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- Crack formation during calcination in air observed, cracks heal and spheres shrink
- Fractures and shrinkage differs significantly for Ce^{IV} compared to Ce^{III} precursor
- Macroscopic behaviour related to release of volatile decomposition products
- Less rapid reduction of α - U_3O_8 matrix to $UO_{2\pm x}$ for Ce^{IV} precursor (Ar:H₂, 700 °C)
- Shrinkage mainly attributed to sintering effects and less to phase transitions

Structural changes of Nd- and Ce-doped ammonium diuranate microspheres during the conversion to $U_{1-y}Ln_yO_{2\pm x}$

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Abstract

The structural changes of ammonium diuranate (ADU) microspheres, prepared by the sol-gel route via internal gelation, were investigated during thermal treatments in oxidative and reducing conditions. In particular, *in-situ* simultaneous thermogravimetric analyses and *in-situ* high-temperature scanning electron microscopy investigations were carried out on un-doped ADU microspheres, and ADU microspheres doped with 10 mol% neodymium or 10 mol% cerium. Calcination in air caused the surface of the particles to crack, but with increasing temperature up to 900 °C some healing, and additionally, shrinkage occurred. The extent of the fractures and the amount of shrinkage was, however, significantly more pronounced in a particle prepared with a tetravalent Ce precursor, as compared to particles prepared with a trivalent Nd or Ce precursor. The macroscopic behaviour could be related to the release of volatile decomposition products and the calcined compositions were identified as (*Ln*-doped) α - U_3O_8 by X-ray powder diffraction. During thermal treatment in reducing conditions a transition from a (*Ln*-doped) α - U_3O_8 phase to a (*Ln*-doped) $UO_{2\pm x}$ phase was observed. After exposure to a hydrogen containing gas mixture, this transition occurred rapidly in un-doped particles, and in particles prepared with trivalent Nd or Ce dopant precursors at 700 °C. Despite the fast reduction reaction, severe fractures appeared in the particles above temperatures of 850 °C, indicating that such behaviour is mainly attributed to sintering effects and less to the phase transition. In contrast, a more delayed reduction reaction was observed in particles prepared with a tetravalent dopant precursor and the described effects appeared less severe.

Keywords: nuclear fuel fabrication, co-conversion, sol-gel, internal gelation, GenIV

1. Introduction

In the context of *minor actinide* (MA) recycling of spent nuclear fuel, research on alternatives to the conventional powder processing for the fabrication of MA-containing transmutation fuel is currently carried out. To avoid handling of fine powders and to facilitate automation for the production of nuclear fuel particles, sol-gel processes have been explored for more than six decades [1–6]. The sol-gel routes via internal- and external gelation are common sol-gel methods, which are based on the precipitation of a metal nitrate solution by ammonia into a metal hydroxide. The major difference between both methods is the source of ammonia. During the external gelation, the sol is dropped via a layer of gaseous ammonia into ammonia solution and the precipitation is triggered by a NH_3 mass transfer into the sol droplet [7]. It occurs while the droplet is falling,

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resulting in a spherical particle. In the internal gelation (IG), a heat transfer causes reactants in the sol to get protonated and decompose thermally to ammonia directly within sol droplets [8]. The resulting pH increase causes the solidification. The spherically shaped precipitate is thermally treated and can afterwards directly be used as particle fuel [9] or compressed into the commonly used fuel pellets [10]. The IG process offers a dust-free fabrication route to convert mixtures of different metal cations into a homogeneous precipitate and was used to prepare MOX particles [11] and U/Pu nitride particles [12]. Moreover, it might be a promising strategy to produce americium-containing transmutation fuel [13], as envisaged by the scientific community to close the nuclear fuel cycle.

The IG process was applied to fabricate Nd-doped $\text{UO}_{2\pm x}$ particles with up to 43 mol% Nd [14], using neodymium as simulant for the trivalent MA americium. It was optimised in terms of automation and its suitability to prepare Ce-doped $\text{UO}_{2\pm x}$ microspheres was investigated using Ce^{III} and Ce^{IV} dopant precursors to simulate plutonium [15]. Microspheres containing up to 30 mol% Ln were fabricated using Nd^{III} and Ce^{III} , while Ce^{IV} was suitable to prepare microspheres with up to 10 mol% dopant, using the chosen gelation parameters. The dopant was successfully incorporated into the $3\text{UO}_3 \cdot 2\text{NH}_3 \cdot 4\text{H}_2\text{O}$ matrix [16] of the dried gels, which was converted into $\text{U}_{1-y}\text{Ln}_y\text{O}_{2\pm x}$ solid solutions by a calcination in air at 900 °C for one hour and a subsequent sintering in reducing atmosphere at 1600 °C for 10 h. Sintering under reducing conditions resulted in broken particles for the compositions prepared with the Ce^{III} precursor and higher dopant contents (≥ 15 mol%), while compositions prepared with Ce^{IV} exhibited cracks on the surface [15]. The thermal decomposition of un-doped ADU particles during the calcination in air was investigated by a combination of thermogravimetric analyses (TGA), thermogravimetric differential scanning calorimetry (TG-DSC), evolved gas analysis mass spectrometry (EGA-MS) and *in-situ* high-temperature X-ray powder diffraction (HT-XRD) [16]. However, no information on the dopants' influence on the thermal decomposition of Ln -doped ADU particles prepared by IG are available in literature, which was the motivation to investigate the thermal treatment of the gelled particles during their conversion into $\text{U}_{1-y}\text{Ln}_y\text{O}_{2\pm x}$. Moreover, we want to study the particle behaviour during the sintering under reducing conditions.

Within this work, the structural changes of un-doped ADU particles, and ADU particles doped with about 10 mol% Nd or Ce were analysed during the conversion to (Ln -doped) $\text{UO}_{2\pm x}$. For the Ce-doped material, particles prepared with a trivalent and a tetravalent dopant precursor were selected. The thermal treatment was performed using simultaneous thermogravimetric analyses (STA) to investigate the decomposition and phase transitions occurring during a calcination in air ($T_{\text{max}} = 900$ °C, 1 h) and a sintering in reducing atmosphere ($T_{\text{max}} = 1000$ °C, 2 h). Moreover, *in-situ* high-temperature scanning electron microscopy (HT-SEM) investigations were carried out, simulating also both thermal treatment steps.

2. Experimental

2.1. Fabrication of particles

The sample material studied within this work was fabricated by the sol-gel route via internal gelation (IG), using acid-deficient uranyl nitrate (ADUN) solution as uranium precursor. The preparation of the ADUN solution ($c(\text{U}) = 2.6$ mol L⁻¹, pH = 1.7, $\rho = 1.85$ g cm⁻³, $\frac{c(\text{NO}_3^-)}{c(\text{U})} = 1.56$) and un-doped ADU gels is described elsewhere [16]. A similar methodology was applied to prepare Nd- and Ce-doped ADU gels with dopant contents up to 30 mol%, using the hexahydrate nitrates of the trivalent lanthanides and diammonium nitrate of cerium(IV) as dopant precursors. The fabrication is described in detail in our previous study [15]. In both studies, urea and hexamethylenetetramine (HMTA) served as gelation agents during the IG

45 process, the gelation agent amount is expressed as ratio (R) of the molar amounts of the gelation agents to the molar amounts of the metals in the sol. The syntheses of all compositions, independent on the dopant, its content and its oxidation state, were carried out with a molar metal concentration of $c(M) = 1.3 \text{ mol L}^{-1}$ and $R = 1.2$ for both gelation agents. Gelation took place in silicone oil ($T = 90 \text{ }^\circ\text{C}$) and the obtained particles were washed thrice with petroleum benzine, aged for 24 h in ammonia solution ($w(\text{NH}_3) = 12.5 \text{ wt}\%$) and washed twice with ammonia solution ($w(\text{NH}_3) = 12.5 \text{ wt}\%$). The un-doped particles and
50 compositions containing about 10 mol% dopant, prepared with Nd^{III} , Ce^{III} and Ce^{IV} , were investigated within this study.

2.2. Thermogravimetric differential scanning calorimetry

TG-DSC analyses were performed using a NETZSCH STA 449 F1 Jupiter thermobalance, having an absolute uncertainty on mass readout including drift and noise of $\pm 0.03 \text{ mg}$ (2σ). The measurements were carried out in Pt/Rh crucibles with heating rates of $10 \text{ }^\circ\text{C min}^{-1}$ and sample masses of about 20 mg. The particles were ground in a mortar prior to the analyses.

55 The set-up was evacuated to $\leq 0.63 \text{ mbar}$, then synthetic air was introduced into the furnace and a measurement was carried out purging the system with 80 mL min^{-1} of synthetic air and the balance compartment with 20 mL min^{-1} of argon as protective gas. After an isothermal equilibration plateau at $40 \text{ }^\circ\text{C}$ (30 min), the calcination in synthetic air was simulated using a heating rate of $10 \text{ }^\circ\text{C min}^{-1}$ and a holding time of one hour at $900 \text{ }^\circ\text{C}$. Afterwards, it was cooled to $40 \text{ }^\circ\text{C}$ ($10 \text{ }^\circ\text{C min}^{-1}$), followed by an isothermal segment at $40 \text{ }^\circ\text{C}$ (60 min) in which the atmosphere was changed from synthetic air to Ar. To simulate the
60 sintering in reducing atmosphere, the samples were heated up to $700 \text{ }^\circ\text{C}$ ($10 \text{ }^\circ\text{C min}^{-1}$) and during an isothermal segment of 60 min at $700 \text{ }^\circ\text{C}$, the gas atmosphere was changed to Ar: H_2 (96:4). The temperature was held for about 30 min before it was further heated to $1000 \text{ }^\circ\text{C}$ at $10 \text{ }^\circ\text{C min}^{-1}$. The temperature was not further increased due to experimental restrictions. An isotherm of 2.5 h was maintained at $1000 \text{ }^\circ\text{C}$ and 30 min prior to its end, the atmosphere was switched back to Ar. Then, the samples were cooled down to $40 \text{ }^\circ\text{C}$ ($10 \text{ }^\circ\text{C min}^{-1}$) and a final isothermal equilibration plateau was maintained for one hour.
65 All gases had a purity of 99.9992 % and originated from AIR PRODUCTS SA. Flow rates of 80 mL min^{-1} were used. The applied temperature program is visually summarised in figure 2a.

The sample mass at the end of the initial equilibration plateau was considered as initial *in-situ* mass. The tare mass of the crucible, as well as the mass of the crucible containing the sample was measured prior to and after the TG-DSC analysis (Mettler-Toledo AT201), having an uncertainty of $\pm 0.04 \text{ mg}$ (2σ). The *ex-situ* masses were used to calculate absolute masses
70 and quantify mass losses occurred during the evacuation of the set-up, since the latter are not logged by the device.

2.3. In-situ high-temperature scanning electron microscopy

The HT-SEM investigations were carried out using a FEI Quanta 200 ESEM FEG environmental scanning electron microscope, equipped with a FEI HT 1400 furnace. Measurements were performed in a custom made crucible with an integrated thermocouple. The sample is located directly on the thermocouples' head, allowing precise temperature measurements. De-
75 tails of the set-up were published by Podor et al. [17].

The sample itself was prepared as follows: a single microsphere was isolated and stricken onto a high temperature embedding ceramic paste (Durapot type from Final Materials). This paste does not react chemically with the microspheres, even at high temperature. The microsphere was positioned in order to remain stable in a hole during the entire heat treatment while minimising the volume of the total montage. The montage was positioned on the head of the thermocouple, and operating
80 conditions were adjusted in order to obtain the best balance between reactivity, sample sensitivity to the electron beam and

image quality. Accelerating voltage was fixed to 20 kV, beam current was adjusted to approximately 0.1 nA and working distance was 20 mm. Imaging conditions were managed by adjusting the bias of the heat shield and contrast, brightness and polarisation of the Gaseous Secondary Electron Detector (GSED) [18].

The thermal treatment during the HT-SEM analyses varied from the one applied in the TG-DSC measurements. Instead of argon, nitrogen was used as inert gas, while a reducing atmosphere was maintained by a N₂:H₂ mixture (95:5). Moreover, the chamber of the device was once filled with the desired atmosphere and no dynamic flushing was carried out. During atmosphere changes, a series of 8 purges with pressures ≤ 5 mbar was performed. The temperature profile was comparable to the one of the TG-DSC analyses. Heating rates of 10 °C min⁻¹ were used as well, with one exception. The heating to 700 °C within the 2nd step, prior to the introduction of a reducing atmosphere, was performed with a heating rate of 30 °C min⁻¹, N₂:H₂ was introduced immediately after reaching 700 °C and the heating was directly continued afterwards. Moreover, a maximum temperature of 1100 °C was applied during the sintering in reducing atmosphere. No active cooling was employed for the cooling processes and the reducing conditions were maintained during the cooling of the last treatment step. During the progress of the temperature profile, images were recorded using different magnifications.

Two distinct image processing strategies were developed to determine the volumetric sample variations from the low magnification images. (1) For the un-doped material, as low magnification images recorded did not allow observing the complete microsphere, several features of the sample surface were tracked automatically along the complete image series using the *Tracking* plugin implemented in the *Fiji* software package (Version 1.52p) [19]. An average value of the distances from these invariant points was calculated for each image all along the image series. Thus they were correlated to the diameter by several comparison points recorded during the experiment. (2) For the three other image series, very low magnification images were recorded allowing to see the whole microsphere during the heat treatment. Thus, after image segmentation using the *Weka Training Segmentation* implemented in the *Fiji* software, the microsphere diameters were automatically calculated for each image of the image series. The background of the images presented in this study was stripped using the *GNU Image Manipulation Program* (GIMP, Version 2.8.18).

2.4. X-ray powder diffraction

The intermediates after a calcination in air were analysed by X-ray powder diffraction (XRD). A PANalytical X'PertPro diffractometer, utilising a Bragg–Brentano parafocusing geometry in a θ - θ configuration was used. The device is equipped with a copper LFF X-ray tube as radiation source. The incident beam path was foreseen with a combination of a fixed divergence slit, 0.02 rad soller slits and a copper beam mask. A nickel filter was introduced into the diffracted beam path. Detection was done with a position-sensitive detector (PANalytical X'Celerator). Zero point calibration was performed on a sintered, high purity silicon disc, while a sintered alumina disc (NIST Standard Reference Material 1976b) was used for validations. The instrument bias was assessed by lattice parameter refinement of silicon and was found to be smaller than 2×10^5 relative (2σ).

The calcination of the samples prior to the measurements was carried out in Al₂O₃ crucibles, using a heating rate of 1.5 °C min⁻¹ and an isotherm of one hour at 900 °C (Nabertherm, LT 9113/P330). Afterwards, the furnace was cooled to a temperature of 150 °C with a rate of 10 °C min⁻¹ and the crucibles were stored in a desiccator to cool to room temperature. XRD samples were prepared by mixing one particle with ethanol in a mortar. The material was ground and the resulting suspension dropped onto a zero background silicon single crystal holder. After the ethanol evaporated, diffractograms were recorded in the range from 20° to 143° 2θ , using a step size of 0.017° 2θ .

Table 1: Lattice parameters a , b , and c , determined for the orthorhombic (Ln -doped) α - U_3O_8 phase present in the particles after a calcination in air at 900 °C for 1 h, as well as the *goodness of fit* (GOF) and *weighted-profile R-factor* (R_{wp}) of the refinements.

composition	a / nm	b / nm	c / nm	GOF / %	R_{wp} / %
un-doped	0.414 55(1)	1.195 49(1)	0.671 82(1)	4.2	8.4
Nd-doped	0.416 28(1)	1.192 58(2)	0.675 52(2)	3.2	6.8
Ce-doped (Ce ^{III})	0.415 19(1)	1.184 34(2)	0.677 19(2)	2.8	6.7
Ce-doped (Ce ^{IV})	0.415 22(1)	1.184 11(2)	0.676 44(2)	1.9	5.0

Lattice parameters were determined using the Pawley refinement option of the software package *HighScore Plus* by PANanalytical (Version 4.8). A parameter set consisting of a cubic Chebyshev polynomial to model the background, the scale factor and lattice parameters for the (Ln -doped) α - U_3O_8 phase, along with a global parameter for sample displacement and a Pseudo Voigt profile function, was applied during the refinements.

3. Results

The actual dopant contents of the particles were determined by inductive coupled plasma mass spectrometry in our former study [15]. Two particles of each composition, dissolved in HNO_3 , were diluted and analysed. Molar metal fractions of $\chi(Nd) = 11(1)$ mol%, $\chi(Ce) = 10(1)$ mol% (Ce^{III} precursor) and $\chi(Ce) = 10(1)$ mol% (Ce^{IV} precursor) were determined for the particles examined within this work. For simplicity, the aimed molar metal fractions of 10 mol% are used to describe and discuss the results in this study. All uncertainties of this work are given for a 95 % confidence level (2σ).

3.1. XRD analyses

Within this study, intermediate products after the calcination in air at 900 °C were analysed by XRD. Figure 1a shows the XRD pattern obtained for the un-doped material, which agrees to the reference pattern for α - U_3O_8 (PDF-2: 01-072-1078, [20]). For the Nd- and Ce-doped compositions (figure 1b and 1c), a different degree of crystallinity was observed, leading to broader reflections, which do not allow to deconvolute and identify the individual reflections occurring at about 26°, 34°, 44°, 46° and 52° 2θ . For the doped compositions, no additional reflections were observed and a slight shift of the reflections to higher 2θ angles can be noticed, indicating a lattice contraction that might be caused by the incorporation of the dopant into the α - U_3O_8 matrix. The XRD data were used to determine the lattice parameters a , b , and c , for the orthorhombic (Ln -doped) α - U_3O_8 phase, present in the particles after the calcination in air. The results, as well as the *goodness of fit* (GOF) and *weighted-profile R-factor* (R_{wp}) of the refinements, are listed in table 1.

3.2. TG-DSC analyses

The TG-DSC measurements were carried out on 5 to 7 ground particles, ensuring good heat conductivity of the sample material and the crucibles' bottom, which is required to obtain minimal differences between the actual sample temperature and the temperature measured, and to obtain sharp DSC signals. Conversely, the HT-SEM investigations were performed on a single microsphere, which could possibly lead to inconsistent findings between both techniques. To estimate the effect of the samples' physical shape on the analysis, un-doped spheres were additionally measured as-produced (unground) via TG-DSC.

The calcined material was identified for all samples as (Ln -doped) α - U_3O_8 , as described in section 3.1. Furthermore, the un-doped material was identified by *in-situ* HT-XRD as α - U_3O_8 at a temperature of 665 °C [16]. Since no significant mass

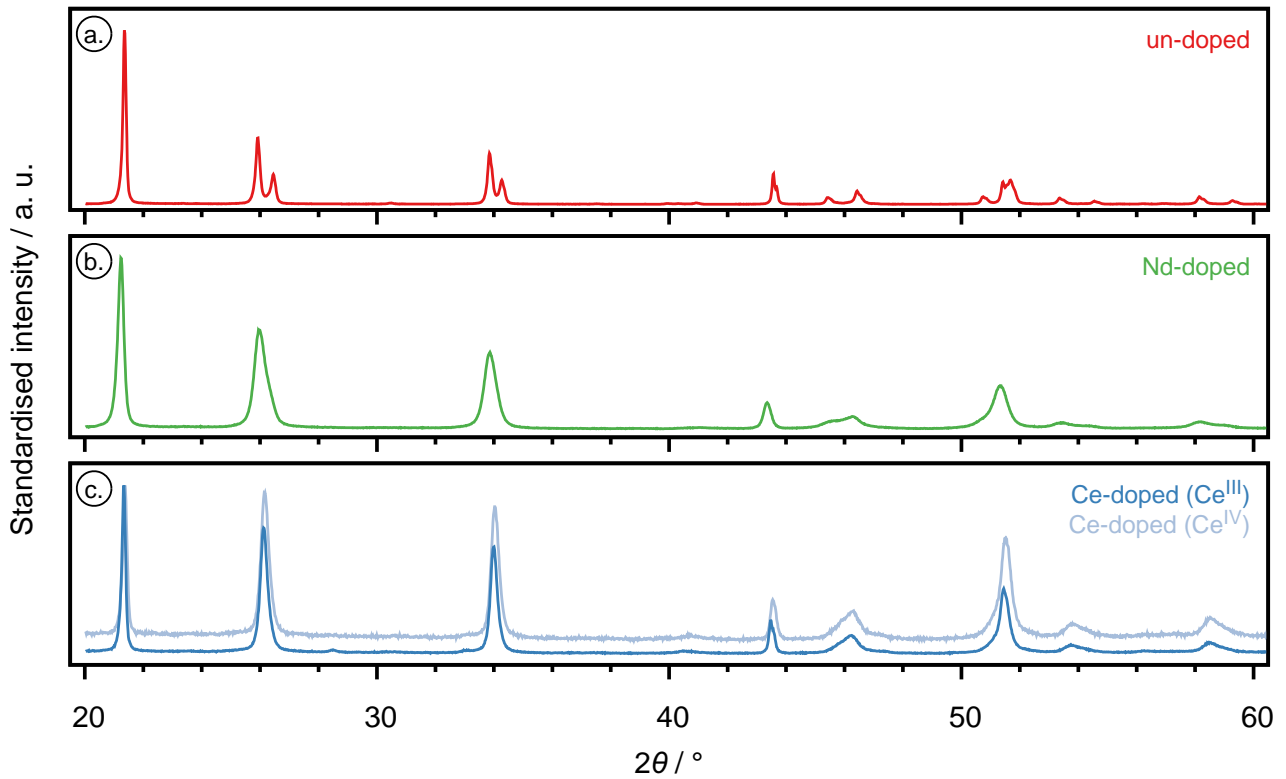


Figure 1: XRD reflection patterns of the un-doped (a), 10 mol% Nd-doped (b) and 10 mol% Ce-doped (c) particles after a calcination in air at 900 °C for 1 h.

Table 2: Initial and final *ex-situ* and *in-situ* masses of the TG-DSC analyses, as well as masses at 83.5 min (550 °C, air), 420 min (700 °C, Ar) and 550 min (1000 °C, Ar:H₂), data are normalised to the sample mass at 665 °C during heating in air, uncertainties are given with a confidence level of 2 σ .

sample		un-doped (as-produced)	un-doped	10 mol% Nd	10 mol% Ce (Ce ^{III})	10 mol% Ce (Ce ^{IV})
$m(\text{initial})_{\text{ex-situ}}$	/ mg	22.12(4)	21.04(4)	18.81(4)	21.89(4)	20.97(4)
$m(\text{initial})_{\text{ex-situ}}$	/ wt%	117.1(2)	116.6(2)	115.8(2)	114.6(2)	120.0(2)
$m(\text{initial})_{\text{in-situ}}$	/ wt%	114.3(3)	112.8(3)	112.1(3)	110.7(3)	116.5(3)
$m(83.5 \text{ min})_{\text{in-situ}}$	/ wt%	102.0(3)	102.7(3)	102.5(3)	102.4(3)	101.9(3)
$m(420 \text{ min})_{\text{in-situ}}$	/ wt%	100.5(3)	100.1(3)	100.0(3)	99.9(3)	100.0(3)
$m(550 \text{ min})_{\text{in-situ}}$	/ wt%	97.0(3)	96.5(3)	96.5(3)	96.2(3)	96.3(3)
$m(\text{final})_{\text{ex-situ}}$	/ wt%	100.6(2)	99.9(2)	98.3(2)	97.2(2)	98.3(2)
$m(\text{final})_{\text{ex-situ}}$	/ mg	19.02(4)	18.04(4)	15.96(4)	18.57(4)	17.18(4)

differences were observed between the sample mass present at 665 °C and the calcination products (as described later on), the obtained data were normalised to the masses present at 665 °C. Figure 2a visualises the temperature program applied during the TG-DSC analyses. The masses at the end of the initial equilibration isotherm were considered as initial *in-situ* masses. The individual values, as well as the initial and final *ex-situ* masses are included in table 2.

150 3.2.1. Comparison of as-produced and ground un-doped ADU microspheres

Six particles of the un-doped composition were measured by TG-DSC with and without out grinding them beforehand. We observed a slightly smaller absolute initial mass for the 6 ground particles compared to the 6 as-produced particles (table 2, $m(\text{initial})_{\text{ex-situ}}$), which can be explained by losses during grinding and variations in particle masses. The differences between the initial *ex-situ* and initial *in-situ* masses correspond to loosely adsorbed water molecules, which were removed during the

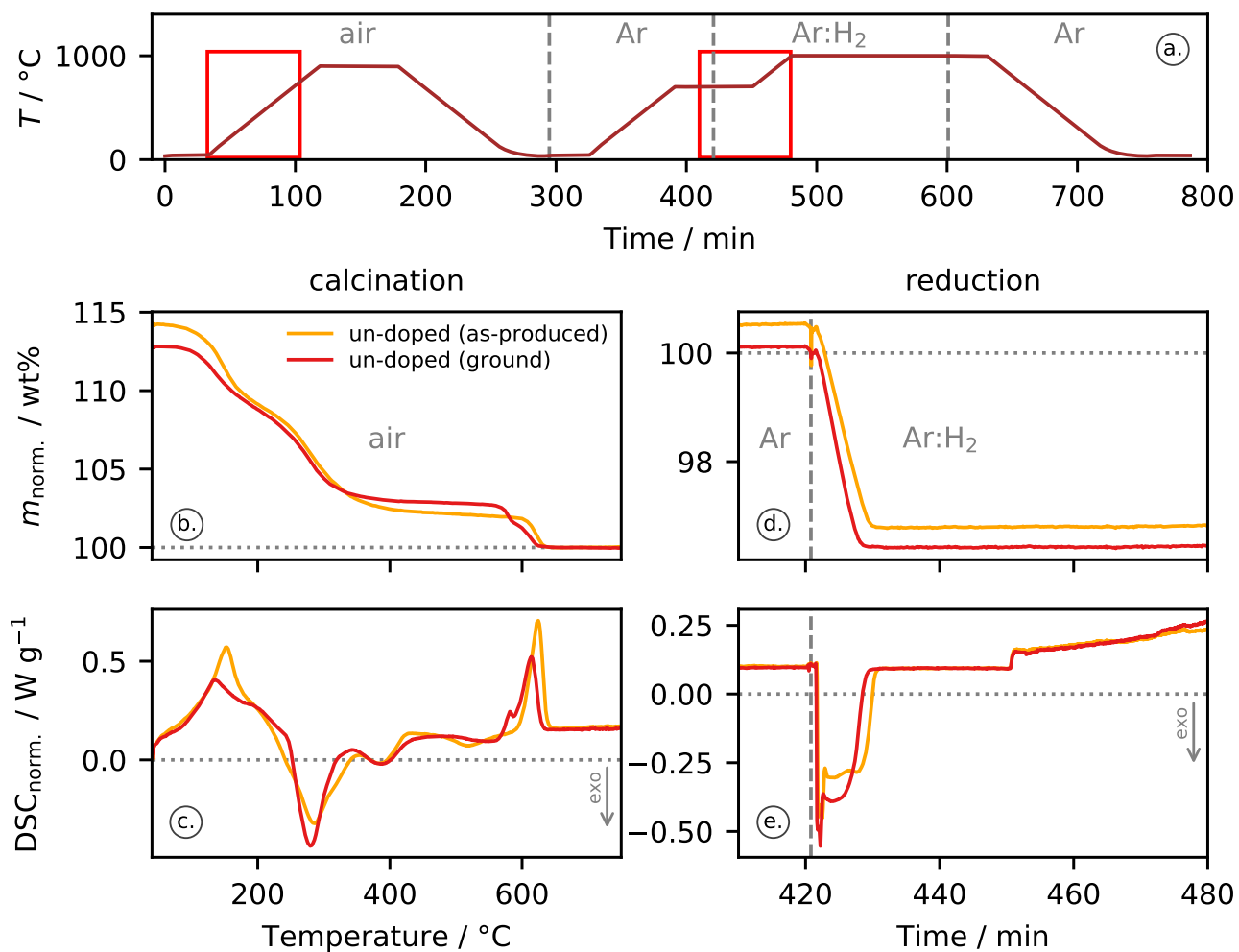


Figure 2: Temperature program applied during the TG-DSC analyses (a), masses and DSC signals obtained for the ground and as-produced, un-doped particles during the heat-up step of the calcination in air as function of temperature (b, c), and after introducing a reducing atmosphere as function of time (d, e), as indicated by the red rectangles in (a). Data are normalised to the masses present at 665 °C.

155 purging of the set-up prior to the measurement. During this step, a mass loss of $-3.7(4)$ wt% was determined for the ground sample and a smaller loss of $-2.8(4)$ wt% for the as-produced sample. The variation is most likely caused by the removal of a larger amount of loosely adsorbed water from the ground particles. Moreover, different normalised initial *ex-situ* masses were observed (table 2), which might be caused by an exchange of NH_3 with H_2O molecules, in this type of ADU, as described by Cordfunke [21] and discussed in our earlier study [16].

160 The normalised masses, for the heat-up segment of the calcination step in air, are plotted as function of the temperature in figure 2b. The corresponding region of the temperature program is indicated in figure 2a by the red rectangle on the left-hand side. During the heating of the calcination in air, comparable masses of about 110 wt% are present in both samples at 166°C . Upon temperature increase, reactions with similar mass losses take place in both samples up to 296°C (figure 2b), mass differences of $-5.3(3)$ wt% and $-5.2(3)$ wt% were determined for the as-produced and ground samples, respectively. A close inspection of the mass signals reveals a mass shift to slightly higher temperatures for the as-produced particles, showing that they were not heated up as fast as the ground material. Due to the smaller surface available for the heat exchange between the sample and the crucibles' bottom, which is located on a stage connected to the thermocouple, a longer time is needed to measure a temperature increase, compared to the ground sample with a larger heat exchange surface. Figure 2c shows the normalised DSC signal for the heating in air. The signals indicate an exothermal reaction for this region, here we see a notably better resolution for the ground sample material, which is a consequence of the described effect.

170 During the further temperature increase from 296°C to 550°C , we observed a continuous mass decrease for the as-produced samples, which is less pronounced for the ground sample (figure 2b). Mass losses of $-3.0(3)$ wt% and $-1.8(3)$ wt% were determined for the as-produced and ground samples, which differ significantly. Between 550°C and 587°C , the as-produced sample shows no noticeable mass loss ($-0.1(3)$ wt%), while we observe a mass loss step with a difference of $-1.1(3)$ wt% for the ground sample, followed by mass loss steps for the conversion to $\alpha\text{-U}_3\text{O}_8$ with mass differences of $-1.9(3)$ wt% and $-1.6(3)$ wt% up to the 665°C . In this case, the as-produced material behaves similar to previous observations made by TGA and TG-DSC on ground samples of the identical particle batch [16]. Moreover, the result matches with the theoretical mass difference of -1.90 wt% for the conversion of UO_3 to $\alpha\text{-U}_3\text{O}_8$ and the conversion occurred in one step, as the U-O phase diagram suggests [22, p. 354, fig. 5.19]. A reaction enthalpy of $71(4)$ kJ mol^{-1} was determined for the as-produced sample, while $65(4)$ kJ mol^{-1} was found for the ground sample, taking both reaction steps into account. The difference between both values might be related to the heat transfer efficiency: Because of their shape, the as-produced particles have a less optimal thermal conductivity with the DSC crucible, which results in lag on the temperature readout and a false impression of additional endothermic behaviour. In our previous study [16], we determined a reaction enthalpy of $72(2)$ kJ mol^{-1} for the un-doped particles in the temperature range between 588°C and 624°C . Although the measurement was performed on ground material of the identical sample batch used in this work, the enthalpy value shows better agreement to the result of the as-produced samples presented here. However, the sample quantity was about a factor two higher in the previous study, which might also have caused some lag on the temperature readout. Within the isothermal segment of the calcination step, the subsequent cooling, and the heating in Ar, we measured a mass increase for both samples. Prior to the introduction of Ar:H_2 a mass increase from 100 wt% present at 95 min (665°C) by $0.5(4)$ wt% was observed for the as-produced sample, and an increase by $0.1(3)$ wt% for the ground sample (table 2, $m(420 \text{ min})_{in-situ}$). While the difference for the ground sample is within the uncertainty, the increase for the as-produced material is notable and visually reflected in figure 2d. However, during the subsequent mass loss step due to the reduction to $\text{UO}_{2\pm x}$, we observed identical differences of $-3.6(4)$ wt% for both samples,

which are smaller than the theoretically expected mass difference of -3.80 wt% for the conversion of α - U_3O_8 to stoichiometric UO_2 , but overlap taking the uncertainty into account. The corresponding DSC signal (figure 2e) shows a slightly longer time until the reaction is completed for the as-produced spheres. Reaction enthalpies were determined as $-160(5)$ kJ mol $^{-1}$ and $-156(5)$ kJ mol $^{-1}$ for the as-produced and ground sample, respectively. They differ slightly, but the uncertainties overlap. The enthalpies determined for both conversions, UO_3 to α - U_3O_8 and α - U_3O_8 to $\text{UO}_{2\pm x}$, shows that there are important differences between experiment and theory: When further comparing the experimental reaction enthalpy values to literature it becomes evident that more dedicated work will be required to accurately probe the enthalpy of the conversion reactions in uranium compounds, which is out of scope of this work.

Comparable masses after the oxidation to α - U_3O_8 and similar mass losses associated with the reduction of α - U_3O_8 to $\text{UO}_{2\pm x}$ for both samples indicate that the mass increase observed for the as-produced particles is related to the sample and might be explained by a movement of the spheres in the crucible, induced by the volume shrinkage of the particles, as demonstrated by the HT-SEM results introduced later on (figure 6c and 6f). The comparison showed that differences are notable during the TG-DSC analyses and sample material in the ground form is more suitable for precise measurements. However, the uncertainties related to the radii measured during the HT-SEM analyses of the particles are larger and have a higher impact on a comparison of the results obtained by both methods. Consequently, we conclude that correlations of the TG-DSC and HT-SEM data are based on causalities.

3.2.2. Comparison of ground microspheres

Initial *in-situ* masses between 110.7(3) wt% and 112.8(3) wt% were measured for the ground samples of the un-doped and doped compositions prepared with the trivalent dopant precursors (table 2). A significantly higher initial *in-situ* mass of 116.5(3) wt% was observed for the compositions prepared with Ce^{IV} as dopant precursor. All ground samples exhibited comparable amounts of loosely adsorbed water molecules, mass differences from $-3.9(4)$ wt% to $-3.5(4)$ wt% were determined (differences of $m(\text{initial})_{\text{in-situ}}$ and $m(\text{initial})_{\text{ex-situ}}$, table 2).

The results of the measurements for all ground samples are visualised in figure 3. The plot contains the data-series of the ground un-doped material, described in section 3.2.1. The normalised masses and DSC signals are summarised for the calcination step in subfigures (a) and (b), and for the sintering in reducing environment in subfigures (c) and (d), the shown regions correspond to those presented in figure 2. Up to a temperature of 550 °C (83.5 min), the release of water and ammonia molecules from the initial compound is expected, leading to a UO_3 composition for the un-doped compound [16]. The *in-situ* masses present at 83.5 min are listed in table 2. Taking the initial *in-situ* masses into account, mass differences between $-8.3(4)$ wt% and $-14.5(5)$ wt% were determined for the release of H_2O and NH_3 during the thermal decomposition of all ground samples (table 2, $m(83.5 \text{ min})_{\text{in-situ}}$). The mass differences observed for the un-doped composition ($-10.1(4)$ wt%) and the compounds prepared with the trivalent dopant precursors (Nd: $-9.6(5)$ wt%, Ce: $-8.3(4)$ wt%) are quite comparable, whereas the compound prepared with Ce^{IV} exhibits a significantly higher mass loss during the heating in air up to 550 °C ($-14.5(5)$ wt%).

The last reaction occurring during the calcination in air is the conversion of UO_3 into α - U_3O_8 . The minimum of the first derivative of the masses (DTG) was used to determine temperatures at which this reaction occurred. For the ground, un-doped material, we observe two steps in the relevant region at temperatures of 580 °C and 611 °C, while for the as-produced and the ground doped compounds this conversion took place in one continuous step, as expected taking the U-O phase diagram into

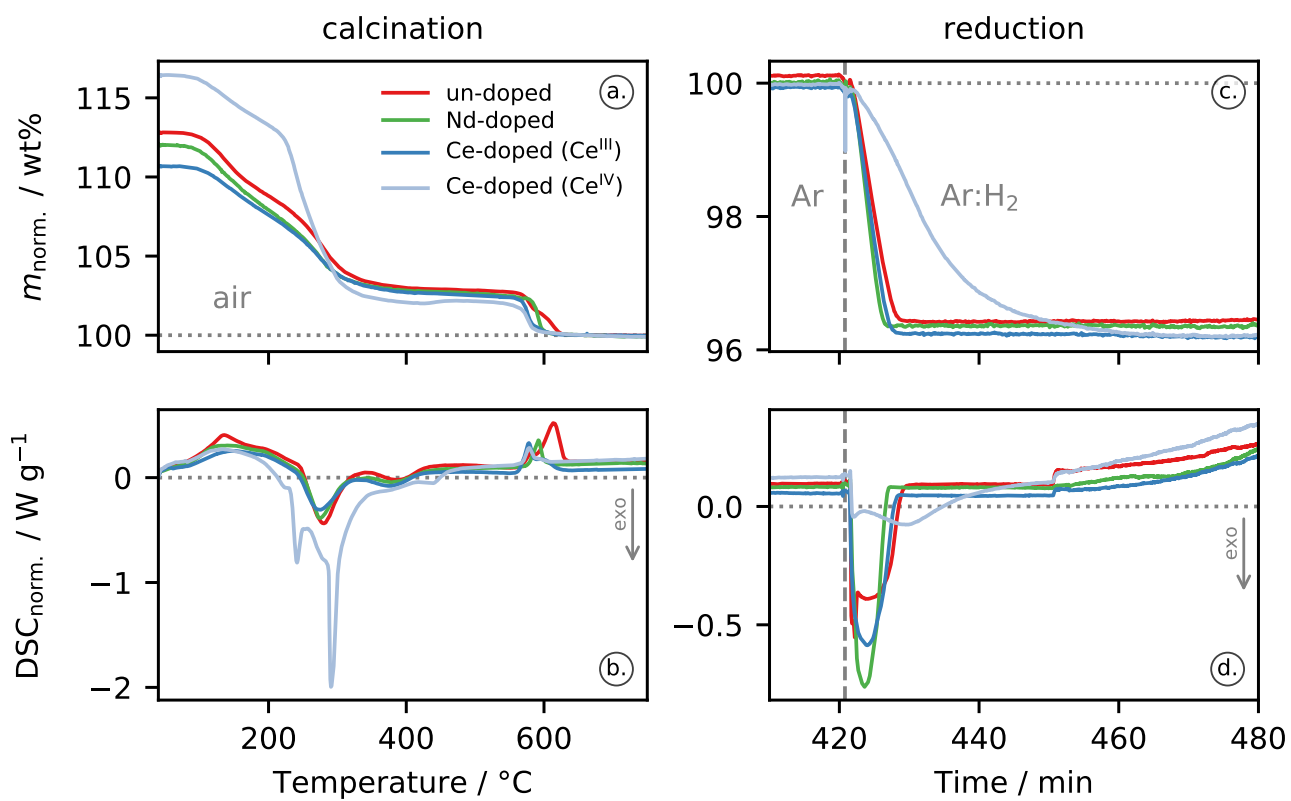


Figure 3: Masses and DSC signals obtained for ground un-doped and Ln -doped particles during the heating of the calcination in air as function of temperature (a, b), as well as the measured data after introducing a reducing atmosphere as function of time (c, d). The masses and DSC data are normalised to the masses present at 665 $^{\circ}\text{C}$. The plots correspond to regions in the temperature program of figure 2a marked with red rectangles.

230 account [22, p. 354, fig. 5.19]. A conversion temperature of 592 °C was determined for the 10 mol% Nd-doped composition, which is in between of the temperatures identified for the two steps in un-doped material, while for the 10 mol% Ce-doped compositions lower conversion temperatures were observed. Comparable temperatures of 576 °C were determined for the material prepared with the Ce^{III} and Ce^{IV} precursors. A stable mass was observed for all compounds at a temperature of 665 °C (95 min). Since the data are normalised to the masses present at 665 °C, the differences between $m(83.5 \text{ min})_{in-situ}$ (table 2) and 100 wt% correspond to the individual mass losses, which can be assigned to this conversion reaction (−1.9(5) wt% to −2.7(4) wt%).

Figure 3b shows the normalised DSC signal for the heating in air. The un-doped material and the doped compositions prepared with the trivalent dopant precursors show the same trends, with slightly varying intensities, while the Ce-doped composition prepared with Ce^{IV} shows a significantly different decomposition behaviour up to about 500 °C. The influence of the dopant on the reaction enthalpy for the conversion of UO₃ to α -U₃O₈ was analysed. For the un-doped material, the heat related to the conversion (integral of the DSC signal) was divided by the molar amount of α -U₃O₈, which was calculated from the mass present at 665 °C, leading to a reaction enthalpy of 65(4) kJ mol^{−1}. For the doped material, we assumed additionally to α -U₃O₈ the presence of Nd₂O₃ or CeO₂, respectively. The actual dopant molar metal fractions were used to calculate molar masses for the compounds present at 665 °C and to determine α -U₃O₈ weight fractions. The corresponding sample masses were multiplied with the α -U₃O₈ weight fractions (94 wt%) to calculate the actual molar amounts of α -U₃O₈ in the compounds at 665 °C. A reaction enthalpy of 21(2) kJ mol^{−1} was determined for the α -U₃O₈ phase of the Nd-doped compositions, while reaction enthalpies of 41(3) kJ mol^{−1} and 16(2) kJ mol^{−1} were determined for the α -U₃O₈ phase present in the Ce-doped material prepared with Ce^{III} and Ce^{IV}, respectively. The smaller heat consumption for the conversion of UO₃ to α -U₃O₈ in the doped material shows that our initial assumption is not valid. Depending on the dopant and its precursors' oxidation state, congruent exothermic reaction(s) occur in parallel to the conversion (i. e. interaction of the Ln-oxide phase with the UO₃ and/or α -U₃O₈ phase), leading to the less pronounced endothermic signal compared to the un-doped composition.

The normalised masses observed during the second step of the thermal treatment are shown in figure 3c, the corresponding region is also indicated with a red rectangle in figure 2a. During the heating to 700 °C in Ar, no reactions are expected and the masses do not significantly vary from the masses observed at 665 °C within the heating of the first thermal treatment step. The *in-situ* masses measured before the introduction of the Ar:H₂ mixture are included in table 2 ($m(420 \text{ min})_{in-situ}$). The masses of all ground samples match within ±0.1 wt% to their standardisation masses, the differences are significantly smaller than the uncertainties (±0.3 wt%). As soon as the Ar:H₂ gas mixture is introduced, we observe a sudden mass loss for the un-doped samples and the doped samples prepared with trivalent dopant precursors. Thus the reduction of α -U₃O₈ to UO_{2±x} is completed within less than 10 min at the isotherm of 700 °C and is not notably influenced by the presence of the dopant. In the case of the Ce-doped material prepared with Ce^{IV}, the reduction occurs much slower and a stable mass was reached at 465 min, which is about 45 min after the introduction of the Ar:H₂ mixture and already within the dynamic segment to reach 1000 °C. The masses present at 550 min (1000 °C, Ar:H₂ atmosphere) were assigned to the (Ln-doped) UO_{2±x} masses and are listed in table 2. Comparable mass losses of −3.6(4) wt%, −3.5(4) wt% and −3.7(5) wt% were observed for the reduction of the un-doped, Nd-doped and Ce-doped (Ce^{III} and Ce^{IV} precursors) compositions, respectively.

265 The normalised DSC signals for the second thermal treatment step are shown in figure 3d and resemble the findings described for the mass signal in the previous paragraph. Exothermal signals were observed for all samples after the introduction of the Ar:H₂ mixture. The smallest time difference until reaching the baseline was observed for the Nd-doped compound

(6.9 min), followed by the Ce-doped compositions (7.4 min, Ce^{III} precursor) and the un-doped material (8.6 min). For the Ce-doped material prepared with the Ce^{IV} precursor, the initial intensity of the baseline is approximately reached at the end of the isothermal segment at 700 °C, but a stable mass was found 44.6 min after the introduction of the Ar:H₂ mixture within the heating segment. Reaction enthalpies for the conversion of α -U₃O₈ to UO_{2±x} were calculated in the same way like it was described for the conversion of UO₃ to α -U₃O₈. We determined -156(5) kJ mol⁻¹ for the un-doped material, -164(6) kJ mol⁻¹ for the α -U₃O₈ phase in the Nd-doped sample, and values of -161(6) kJ mol⁻¹ and -154(6) kJ mol⁻¹ for the α -U₃O₈ phase in the Ce-doped material prepared with Ce^{III} and Ce^{IV} precursors, respectively. Due to the relatively small masses present after the reduction, the enthalpies exhibit quite large uncertainties, with an overlapping region. This leads to the conclusion that the dopant does not significantly influence the reduction of α -U₃O₈ to UO_{2±x}. At a time of about 450 min, an increase in the DSC signal is observed for all samples, which is caused by the change from isothermal conditions to the dynamic heating segment. It does not require further interpretation because it is an experimental artefact.

After an exchange of the atmosphere from Ar:H₂ to Ar for the cooling to room temperature we noticed a mass increase, which corresponds for the un-doped material to the mass loss observed during the reduction. This indicates that the sample re-oxidised to α -U₃O₈ within this step (table 2). The oxygen impurity levels in the used gases are negligible, but a limited amount of oxygen leaks into the experimental device (\approx 15 ppm O₂ in 100 mL min⁻¹ Ar). Given the very small quantities of sample used (\approx 19 mg), even such limited amounts of O₂ can result in the re-oxidation of UO₂ in the given time-frame. Therefore, data recorded in the Ar atmosphere at the end of the treatment (figure 2a) are not discussed in this work.

3.3. HT-SEM analyses

Micrographs of the particles recorded during the HT-SEM investigations are summarised in figure 4 for the calcination in air. As experimental parameters were optimised during the study of the un-doped particle, only few images were recorded at low magnification compared to the doped particles. The left column of figure 4 shows one particle of each composition at its initial state at room temperature. The 2nd and 3rd column show the same particles during the heating, and the right column towards the end of the calcinations' isothermal plateau at the maximum temperature of about 900 °C. In the case of the un-doped material, the initial image was taken at an elevated temperature (figure 4a) and the final one during the cooling to room temperature (figure 4d), they were added to figure 4 due to a comparable magnification to the doped particles. The corresponding temperatures and determined radii, normalised to the radii measured on the initial image (figure 4, left column), are shown as well. The absolute initial radii were determined as 851.2 μ m, 626.3 μ m, 726.5 μ m and 671.5 μ m for the un-doped and Ln-doped ADU microspheres prepared with the Nd^{III}, Ce^{III} and Ce^{IV} dopant precursors, respectively.

Comparable particle shrinkages were found during the calcination in air. Normalised radii ranging between 0.81 and 0.86 were measured at the end of the calcination. The largest difference occurred for the Ce-doped particle prepared with Ce^{IV} and the smallest for the Ce-doped particle prepared with Ce^{III}. The relative radii measured for the un-doped and Nd-doped composition, were both found to be 0.82 at the end of the calcination. Even though the relative radii differences during the entire thermal treatment step were comparable, the Ce-doped particle prepared with Ce^{IV} showed a significantly different behaviour during the heating to 550 °C (figure 4, second column). The formation of cracks on the particle surfaces were observed for all compositions, but for the latter it was much more emphasised and went along with a remarkable shrinkage. The normalised particle radii of the heating to 550 °C within the calcination in air are plotted in figure 5 as function of temperature, the normalised radii after the calcination are included as arrows.

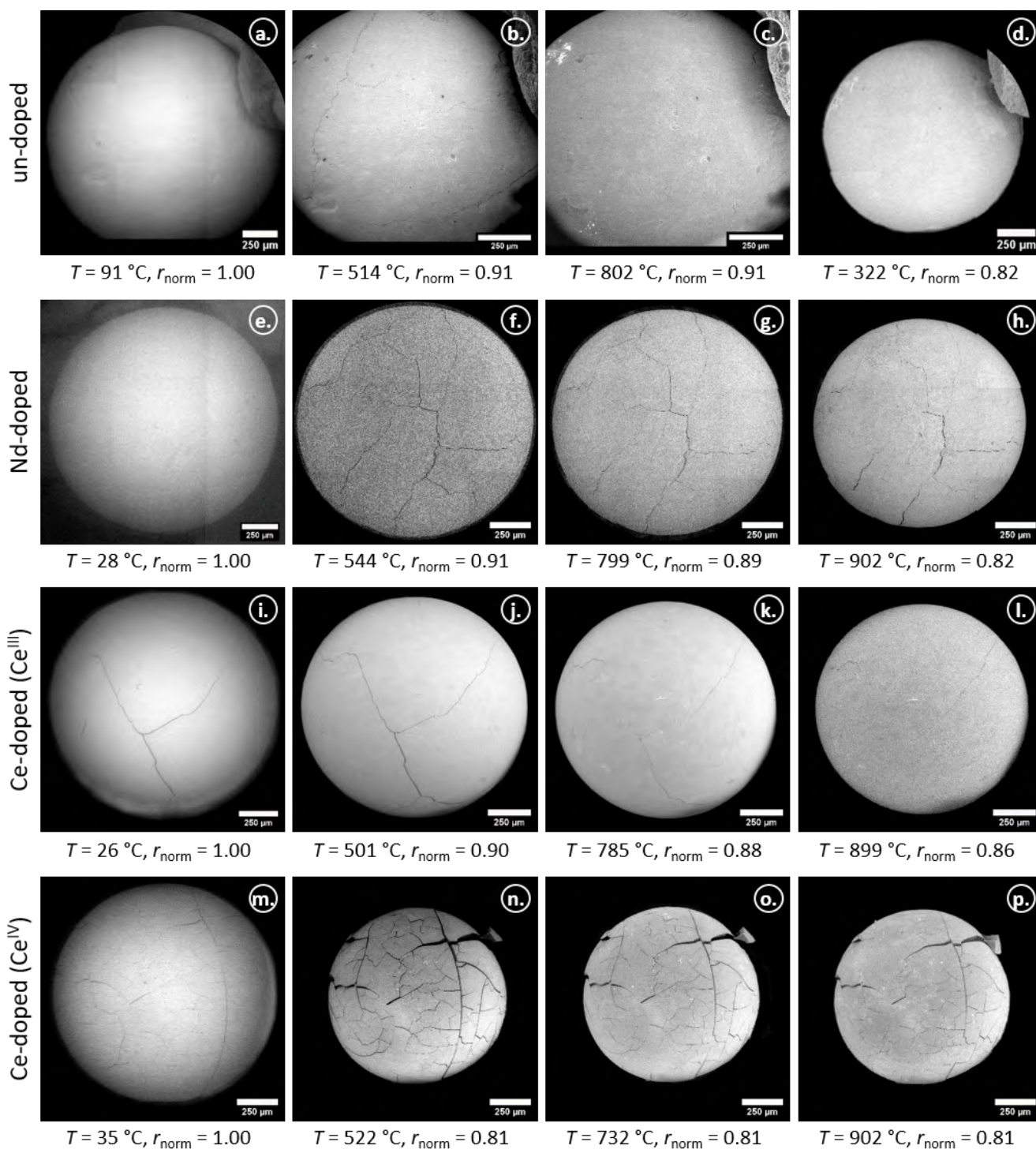


Figure 4: Micrographs of the un-doped and doped particles during the calcination in air. The left column shows one particle of each composition at its initial state, while column 2 and 3 show the same particles during the heating. The images of the right column show the particles towards the end of the isothermal plateau at the maximum temperature of about 900 °C, with an exception for the un-doped particle (taken during cooling at 332 °C, please note also the larger magnifications that had to be used for the un-doped material).

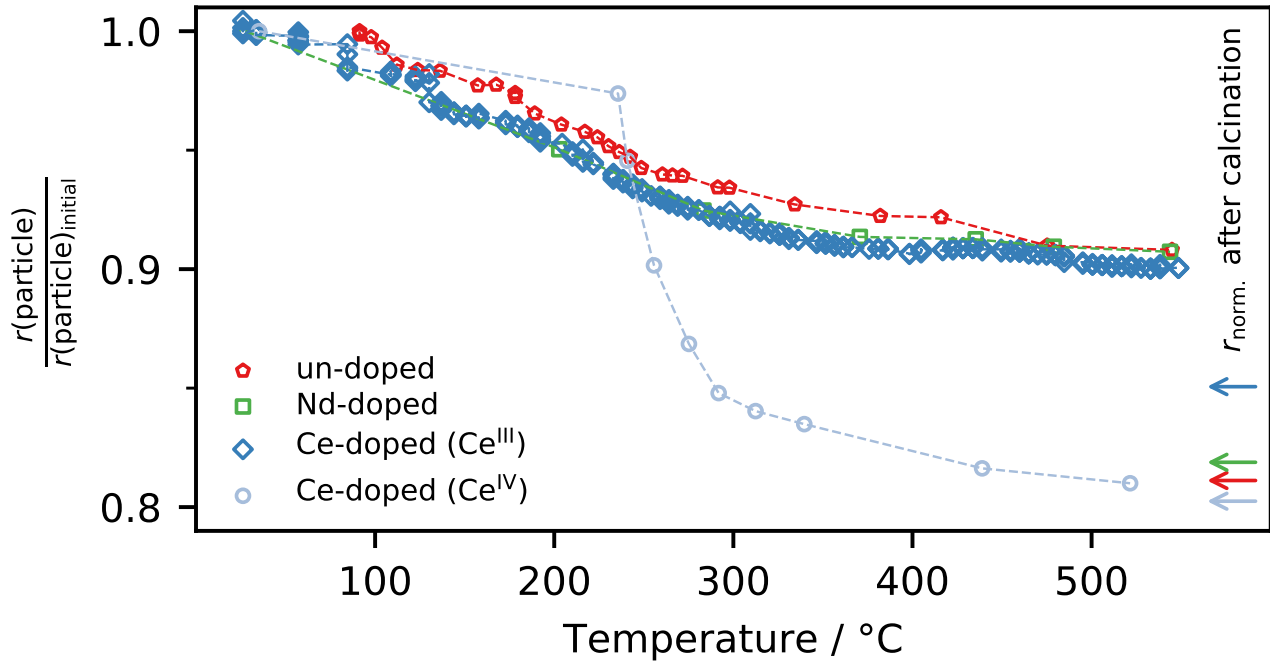


Figure 5: Normalised particle radii during the heating to 550 °C within the calcination in air as function of the temperature (temperature profile shown in figure 6a), as well as the normalised particle radii after the calcination in air (arrows).

305 While all other compositions have a relative radius of about 0.9 between 500 °C and 550 °C, the radius of the Ce-doped particle (Ce^{IV}) shrunk already to a normalised radius of 0.81 (figure 5 and figure 4n), with the most emphasised shrinkage occurring below 350 °C. These results match well with the findings of the TG-DSC analyses, where we observed significant mass losses going along with strong exothermal reactions for the Ce-doped particle prepared with Ce^{IV} in the same temperature region (figure 3b and 3c).

310 The data series obtained from the HT-SEM investigations were split into individual series for the two different thermal treatment steps. Each was offset corrected to match the temperature profile applied during the TG-DSC analyses and to allow a comparison of the data as a function of time. For the calcination in air, the heat up segment to 900 °C was used for this purpose, while the sintering was aligned to the time when a reducing atmosphere was introduced. The temperatures of the individual HT-SEM measurements are shown as a function of the offset corrected times in figure 6a and 6d, for the calcination in air and sintering under reducing conditions, respectively. We observe deviations for temperatures lower than 450 °C and times lower than 60 min within the calcination step, but those results were already described in figure 5 as a function of temperature. The offset corrected data for higher temperatures (T_{max} : 902 °C to 905 °C) and times agree well to each other.

320 In the case of the sintering under reducing conditions, the deviations prior the introduction of the $\text{N}_2:\text{H}_2$ mixture (421 min) are negligible, since no significant expansion or contraction was observed during the cooling to room temperature in air and subsequent heating to 700 °C in N_2 . For the heating process in the reducing environment, the profiles are comparable besides slightly different isothermal holding times at the maximum temperatures (1112 °C to 1131 °C). The *in-situ* particle radii, normalised to the initial radii, measured during both treatment steps within the HT-SEM analyses are plotted in figure 6b and 6e. The development of the particle volumes, calculated from the radii assuming a spherical geometry, normalised to the final particle volumes are shown in figure 6c and 6f.

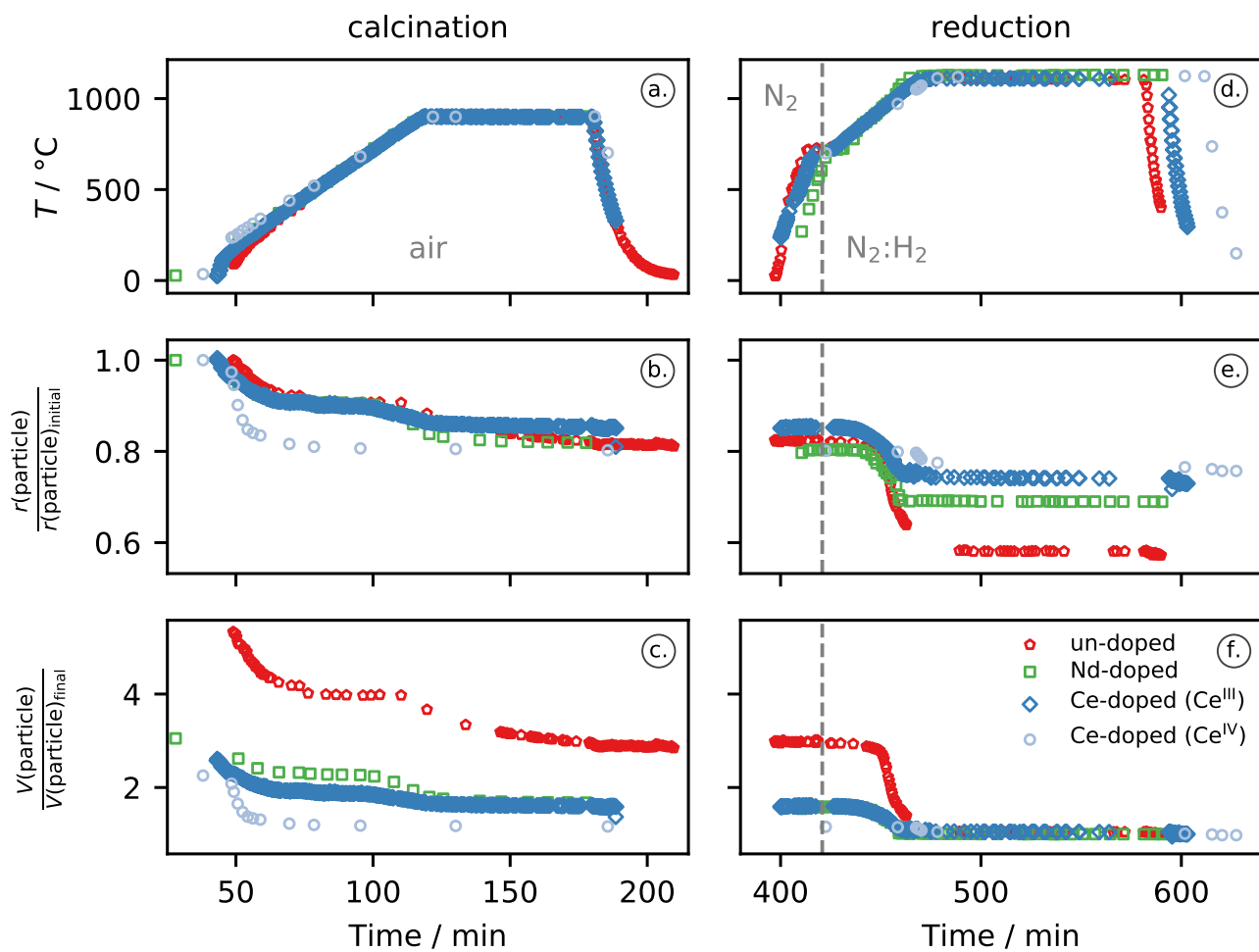


Figure 6: Temperatures measured for the calcination (a) and reduction (d) during the HT-SEM analyses, *in-situ* particle radii normalised to initial radii (b, e), as well as the resulting particle volumes normalised to the final volumes (c, f) as function of the offset corrected time.

325 During the calcination in air, we observe a second significant shrinkage step at temperatures above 700 °C (figure 6b and 6c, about 100 min), which is for all compositions smaller than the shrinkage observed up to 550 °C. The micrographs taken during this step (figure 4, third column) show besides the particle shrinkage that the cracks and gaps, formerly noticed, are less emphasised and closed for the un-doped material.

330 The results for the reduction of the particles in the N₂:H₂ gas mixture (figure 6e and 6f) show prior to the change of the atmosphere to reducing conditions comparable normalised radii to those observed towards the end of the calcination (figure 6b and 6e). Micrographs of the particles during the sintering in the reducing environment are summarised in figure 7, the left column shows images recorded after the introduction of the N₂:H₂ gas mixture at temperatures between 700 °C and 770 °C, the particle radii determined agree well to the radii of the calcined particles. During the treatment in the reducing environment, a shrinkage from normalised radii between 0.80 and 0.85 (figure 7, left column) to radii ranging between 0.57 and 0.77 (figure 7, right column) was observed. The largest difference was measured for the un-doped compound (0.25), while the smallest shrinkage occurred for the Ce-doped composition prepared with Ce^{IV} (0.03). For the doped compositions prepared with the trivalent dopant precursors, comparable differences of 0.11 were observed. Normalised radii differences of up to 0.02 during the further process of the sintering (figure 6e) indicate no remarkable shrinkage after 500 min, which corresponds to an N₂:H₂ exposure time of about 80 min at a temperature of 1112 °C, with an exception for the Nd-doped composition where a slightly higher maximum temperature was applied (1128 °C). The comparison of the individual particles to each other is better reflected by the particle volume normalised to the final volume, as shown in figure 6f.

4. Discussion

345 Samples of the dried gels of the current work were characterised in our previous studies by XRD. The diffraction pattern of the un-doped composition was compared to reference pattern [23] and identified as 3UO₃·2NH₃·4H₂O [16], using the described preparation conditions. A comparable XRD diffraction pattern was observed for the 10 mol% Nd doped compound. Intensity variations were observed, but no additional reflections compared to the un-doped composition were recognised [15]. A significantly different degree of crystallinity was observed for the different Ce precursors. Ce^{III} led to patterns comparable to the Nd-doped particles, while a large fraction of an amorphous phase appeared to be present in the compositions prepared with Ce^{IV}. However, also no additional reflections compared to the un-doped particles were observed for the Ce-doped material [15], allowing us to conclude that the dopant is either incorporated into the ADU matrix, or present as an amorphous phase. Therefore, we assume that the initial sample material of this study corresponds to ADU with 3UO₃·2NH₃·4H₂O stoichiometry, or Nd- or Ce-doped 3UO₃·2NH₃·4H₂O, respectively.

4.1. Structural changes during calcination in air

355 The comparable differences between the initial *ex-situ* and initial *in-situ* masses observed during the TG-DSC measurements (−3.9(4) wt% to −3.5(4) wt%) point to a similar amount of loosely adsorbed water molecules in the initial (*Ln*-doped) 3UO₃·2NH₃·4H₂O material.

360 During the thermal decomposition of the compounds in air, three major mass loss steps were recognised. The first two can be assigned to the removal of water and ammonia [16]. The most significant mass drops occurred at temperatures of about 135 °C going along with an endothermic DSC signal, and at about 275 °C, accompanied with one exothermic reaction for the un-doped material and the doped particles prepared using the trivalent precursors (figure 3a and 3b). For the doped

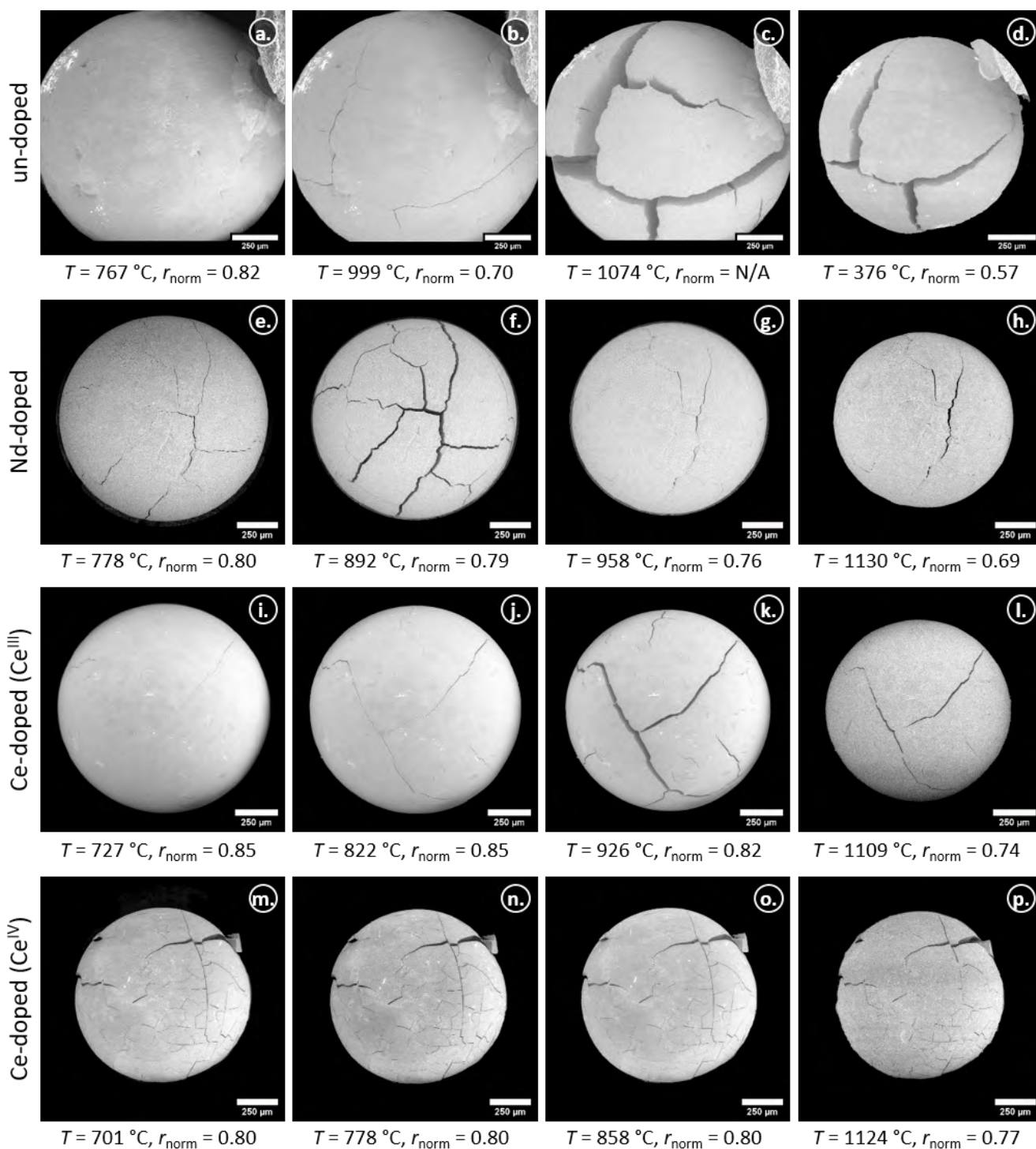


Figure 7: Micrographs of the un-doped and doped particles during the sintering in reducing atmosphere. The left row shows one particle of each composition after the introduction of the $\text{N}_2:\text{H}_2$ gas mixture. The 2nd and 3rd rows show the same particles during the heating to the maximum temperature and the right column shows the particles towards the end of the isothermal plateau at the maximum temperature of about 1100 °C, with an exception for the un-doped particle (taken during cooling at 376 °C, please note also the larger magnifications that had to be used for the un-doped material).

composition prepared with Ce^{IV} , the DSC signal indicates at least four exothermic reactions for the second step, occurring at temperatures of 241 °C, 259 °C, 275 °C and 291 °C, and result in a more pronounced mass loss during the second major mass loss step. The observations match to the thermal decomposition profile of HMTA/urea mixtures [24, 25], which points to residuals of the gelation agents in the particle and an ineffective washing. Similar findings were made for yttria-stabilised zirconia(-ceria) microspheres [24, 26]. Amorphous phases and low porosity prevent the volatile decomposition products to escape the particles and lead to crack formation [27]. An increased HMTA content in the sol resulted in more pore formation, leading to less severe cracking [28]. Moreover, it was demonstrated that the remaining organics can be removed from the particles by a pressurised water treatment of the gelled spheres [26]. Katalenich [25] applied the latter successfully for ceria microspheres, also prepared with diammonium cerium(IV) nitrate as precursor. TGA of the dried residue from the water treatment showed about 90 wt% mass loss up to 400 °C, supporting the assumption of an ineffective washing for the Ce doped batch prepared with the Ce^{IV} dopant precursor.

Close-ups of the particle surfaces are presented in figure 8. Micrographs taken at the beginning of the calcination in air show the initial state (left column) and the same region of the surface at about 900 °C towards the end of the isotherm is presented in the second column. We observe agglomerates for all compositions, the micrograph taken for the composition prepared with Ce^{IV} indicates a poor crystallinity (figure 8m). The latter agrees with the XRD patterns of the initial material, where a large fraction of amorphous material was observed for the Ce-doped composition prepared with Ce^{IV} [15]. It further supports the explanation for the crack formation, volatile decomposition products of the gelation agents are hindered to escape the particle, which results in the fracturing.

The different decomposition behaviour of the composition prepared with Ce^{IV} , compared to all other examined compounds, was also confirmed by the particle shrinkage. The other compositions have a relative radius of 0.9 between 500 °C and 550 °C (figure 5), whereas the radius of the Ce-doped particle prepared with Ce^{IV} shrunk already to a normalised radius of 0.81, with the most pronounced shrinkage below 350 °C. It agrees to the observed mass difference and supports the assumption of retaining gelation agents. Clearly, microspheres prepared by the sol-gel method via IG with a Ce^{IV} dopant precursor have a different crystallinity and/or initial composition than particles prepared with a Ce^{III} dopant precursor. It is most likely a consequence of the diammonium cerium nitrates' $\frac{n(\text{NO}_3^-)}{n(\text{Ce})}$ ratio of 6 as compared to 3, which applies for the trivalent lanthanide dopant precursors and affects the crystallinity [27].

Moreover, we observe a correlation of the decomposition temperatures with the gap development of initially formed cracks. A summary for the un-doped particle is given in figure 9, the closing and opening observed within the first 5 subfigures corresponds to the H_2O and NH_3 release. The gap opening of the third subfigure was observed at 178 °C, and can be assigned to the first major mass loss step observed by TG-DSC. During further heating, this gap closes again. The second major mass loss step started at the temperature of the fourth subfigure (figure 9, 236 °C), but the gap closes further and is not notable at 422 °C anymore. A similar behaviour was observed for the doped particles, prepared with the trivalent Ln precursors.

Comparable masses at a temperature of 550 °C (table 2, $m(83.5 \text{ min})_{\text{in-situ}}$) indicate similar compounds prior to the conversion of the UO_3 matrix to $\alpha\text{-U}_3\text{O}_8$ [16], which is the third major mass loss step during the calcination of the particles in air. The differences caused by the conversion are higher than the theoretically expected ones. Theoretical mass differences of -1.90 wt% (un-doped), -1.77 wt% (10 mol% Nd-doped) -1.78 wt% (10 mol% Ce doped) were calculated, assuming solid solution of $\alpha\text{-U}_3\text{O}_8/\text{Nd}_2\text{O}_3$ and $\alpha\text{-U}_3\text{O}_8/\text{CeO}_2$ at a temperature of 665 °C. However, as result of the mentioned inconsistencies for the conversion observed in the ground and as-produced sample material and too many assumptions, no further attention

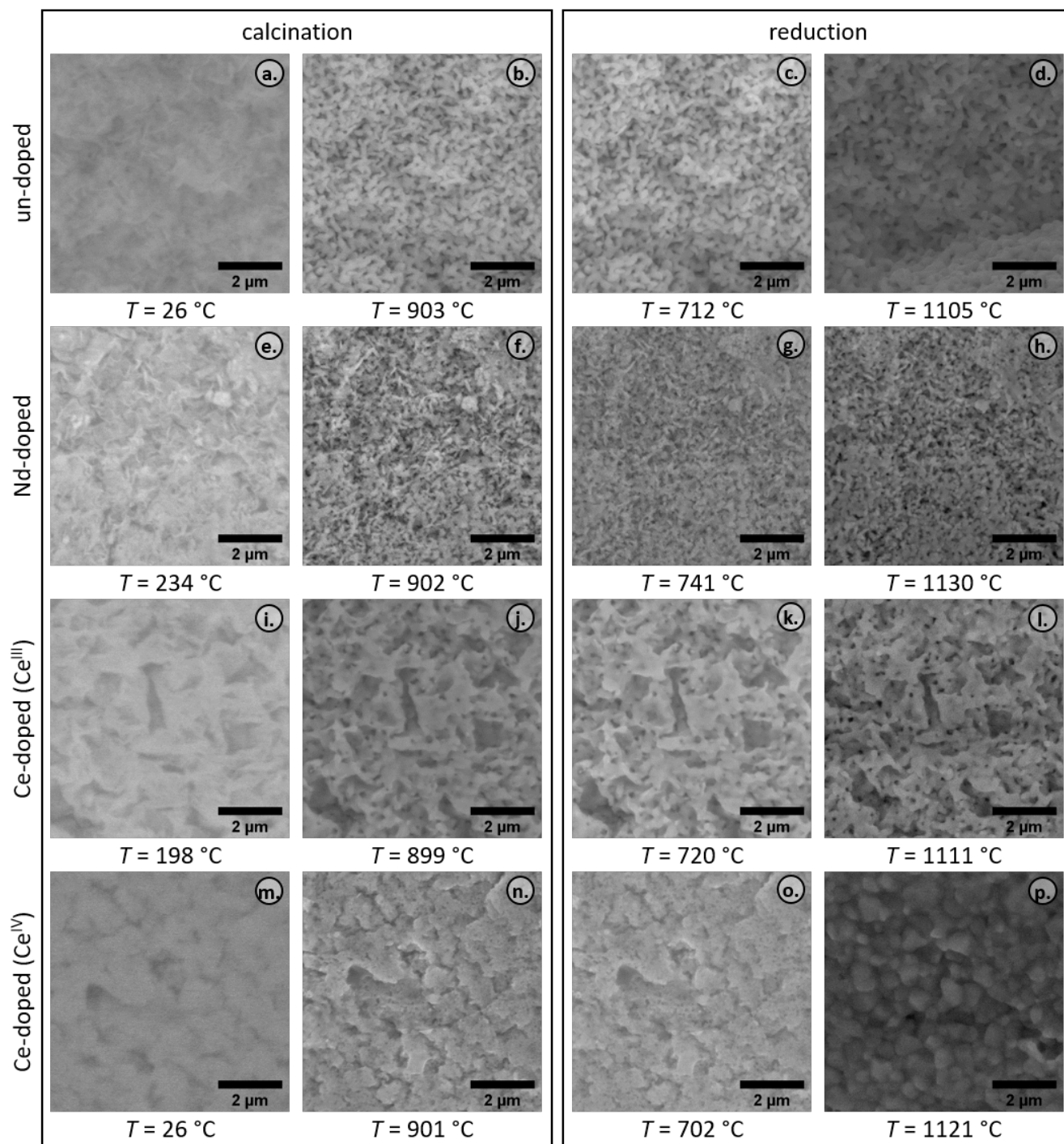


Figure 8: Close-ups of the particle surfaces recorded by HT-SEM during the thermal treatment. The left column shows the particle surface at the beginning of the calcination in air at its initial state, while the 2nd column shows the same region of the surface at T_{max} towards the end of the isotherm. The 3rd column was taken at about 750 °C during the sintering after the introduction of the reducing atmosphere and the last column at about 1100 °C at the end of the sintering isotherm.

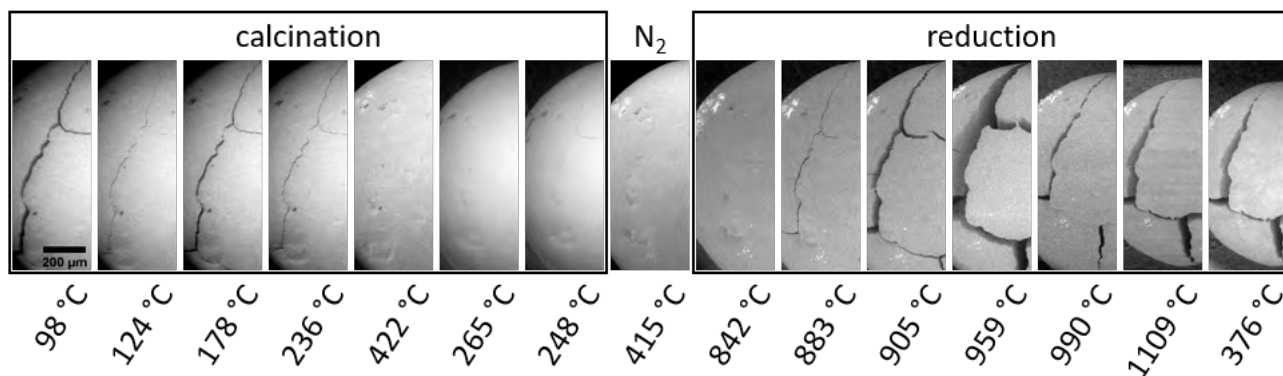


Figure 9: Formation of cracks in the un-doped particle and their development during the calcination in air ($T_{max} \approx 900$ °C) and sintering in reducing atmosphere ($T_{max} \approx 1100$ °C).

is given to this matter and the general decomposition behaviour of un-doped material, described in our former study [16], could be confirmed. The compositions prepared with the trivalent dopant precursor behave similarly, slight mass variations were observed, but the data do not indicate additional reactions during the decomposition in air compared to the un-doped material and the TGA and TG-DSC results of our former study [16]. Since no further gradual mass loss occurred during the calcination in air, the TG-DSC results imply that a treatment temperature of 700 °C is sufficient to form the same product, but the HT-SEM results demonstrated for all particles, besides those prepared with Ce^{IV} , a further significant shrinkage during the heating and the isothermal segment at 900 °C. The shrinkage occurs in the form of a step function for the un-doped material and microspheres prepared with the trivalent dopant precursors, which is most likely caused by the formation of the $\alpha-U_3O_8$ matrix (figure 6b). It is followed by a slow, gradual densification, which is also reflected in the agglomerates. However, it results in quite porous material for the un-doped particle, but individual crystallites can be recognised (figure 8b). The shrinkage of the agglomerates was also observed for the compositions prepared with the trivalent dopant precursors (figure 8f and 8j), both compositions are porous as well. The formation of crystallites can also be seen for the Nd-doped composition, but is less emphasised for the Ce-doped material prepared with Ce^{III} , while less pores were observed for the Ce-doped particle prepared with Ce^{IV} (figure 8n). During the isothermal segment at 900 °C and the subsequent cooling to 265 °C, the gaps of the initially observed cracks were not visible (figure 9). While they were noticed slightly later at a temperature of 248 °C during the cooling.

The presence of comparable compounds after the calcination was confirmed by XRD (figure 1). No additional reflections occurred in the diffraction pattern of the doped compounds, allowing us to assume that the calcination products are solid solution of $\alpha-U_3O_8/Nd_2O_3$ and $\alpha-U_3O_8/CeO_2$. The lattice parameters determined for the un-doped compound are comparable to values published by Loopstra [20], who described the $\alpha-U_3O_8$ phase with the $C2mm$ space group (non-standard setting of space group $Amm2$) and an orthorhombic symmetry. He published lattice parameters of $a = 0.6716$ nm, $b = 1.1960$ nm and $c = 0.4147$ nm. Our lattice parameters (table 1) refer to the standard setting in the $Amm2$ space group, resulting in a swap of the a with c parameter.

4.2. Structural changes during sintering in reducing atmosphere

The introduction of the reducing atmosphere during the thermal treatment caused a quite rapid transition from $\alpha-U_3O_8$ to $UO_{2\pm x}$ at 700 °C (figure 3c and 3d) completed within less than 10 min after exposure to Ar:H₂ (96:4). For the Ce-doped

425 material prepared with the Ce^{IV} precursor, this reduction took about 45 min. The corresponding shrinkage was measured significantly later by HT-SEM (figure 6e and 6f) and occurred during a dynamic segment at elevated temperatures. A comparable behaviour was observed by Clavier et al. for ThO_2 compounds [29]. The mass losses of $-3.6(4)$ wt%, measured in the un-doped sample (ground and as-produced), the smaller mass loss of $-3.5(4)$ wt% of the Nd-doped composition, and the higher difference of $-3.7(5)$ wt% for the Ce-doped material (both precursors), show only marginal differences and are within
430 their uncertainties, but match the trend for the expected mass losses of -3.80 wt%, -3.54 wt% and -3.85 wt%, respectively. The theoretical values were calculated assuming the presence of Nd_2O_3 or CeO_2 prior to the reduction and Nd_2O_3 or Ce_2O_3 afterwards. Reaction enthalpies for the conversion of $\alpha\text{-U}_3\text{O}_8$ to $\text{UO}_{2\pm x}$ were determined as $-156(5)$ kJ mol $^{-1}$ for the un-doped material (as-produced sample: $-160(5)$ kJ mol $^{-1}$), $-164(6)$ kJ mol $^{-1}$ for the Nd-doped sample, and $-161(6)$ kJ mol $^{-1}$ and $-154(6)$ kJ mol $^{-1}$ for the $\alpha\text{-U}_3\text{O}_8$ in the Ce-doped material prepared with the Ce^{III} and Ce^{IV} precursors, respectively. We
435 observe small deviations for the reaction enthalpies, indicating no, or only a marginal, contribution of the dopant to the reaction enthalpy.

For the particles prepared using the trivalent dopant precursors, we observed already comparable shrinkage rates during the isothermal segment of the calcination, if we consider the data standardised to the final particle volumes (figure 6c). Both compounds follow the same trend during the sintering, while the shrinkage of the un-doped material is more emphasised.
440 The initially observed cracks of the un-doped particle (figure 9) became more pronounced during the heating in reducing atmosphere and show a maximum distance between the fragments at about 959 °C, which decreased during further heating. However, the partially open cracks remain present during the subsequent shrinkage and the cooling to room temperature. The significantly smaller shrinkage rate of the Ce-doped particle with Ce^{IV} prepared during the reaction might be explained by its fragmented structure caused by the rapid significant shrinkage during the exothermic reactions within the calcination.
445 Moreover, we showed that less pores were present already after the calcination (figure 8n), compared to the other compounds, which might hinder a H_2 diffusion into the microsphere and caused the slower reduction. Figure 8 includes also the particle surfaces at about 750 °C during the sintering after the introduction of the reducing atmosphere (third column) and at the end of the sintering isotherm at about 1100 °C (right column). The trend observed during the calcination remained for the sintering, while the un-doped and doped particles prepared with trivalent dopant precursors exhibit a large amount of pores, we see a quite dense surface for the composition prepared with Ce^{IV} . Grain growth was observed for all compositions, but significantly larger grains were formed for the Ce-doped composition prepared with the tetravalent dopant precursor.

The *in-situ* behaviour of the particles during the thermal treatment applied within this study is summarised in figure 10, correlating the HT-SEM results with the TG-DSC results. The subplots (a) and (e) show the temperatures of the TG-DSC and HT-SEM analyses as function of the (offset) corrected times for the calcination, and the reduction, respectively. The
455 masses obtained during the TG-DSC measurements were normalised to average masses of dried particles, determined within our previous study [15, 30], and presented in figure 10b and 10f. In the same study average particle radii and volumes were determined, which we used to normalise the radii measured in the HT-SEM investigations and to calculate average particles volumes, assuming a spherical shape (figure 10c and 10g). Both series were combined to determine average *in-situ* particle densities for both thermal treatment steps, which are presented in figure 10d and 10h. It should be noted that the data contain
460 some uncertainties, especially for temperatures below 450 °C and times < 60 min, and for the un-doped composition, where the initial HT-SEM image was recorded at 91 °C. Figure 10 is intended to represent the tendencial behaviour for an average particle of each of the compositions during thermal treatment, and shows that the contribution of the mass changes to the

density is negligible. The particles' density is mainly impacted by the shrinkage.

Volume shrinkage factors normalised to the final volume of 5.3, 3.0, 2.6 and 2.3 for the un-doped, Nd-doped, Ce-doped (Ce^{III}) and Ce-doped (Ce^{IV}) particles were determined after the treatment, carried out in this study (figure 6c and 6f). The results do not match to volume shrinkage factors determined for particles from identical batches after a sintering for 10 h at 1600 °C in Ar:H₂ with an addition of Ar:O₂, corresponding to an oxygen potential of about -420 kJ mol⁻¹, where values of about 7.5, 8.0, 5.7 and 3.0 were found, respectively [15]. The significantly larger volume shrinkage factors show that the investigations carried out in this work describe the structural changes of Nd- and Ce-doped ADU microspheres during the conversion to U_{1-y}Ln_yO_{2±x} very detailed for temperatures up to about 1100 °C, but obviously, sintering is far from being complete at 1100 °C. The presence of U_{1-y}Ln_yO_{2±x} solid solutions was demonstrated by XRD for all doped compositions analysed in this work, after the sintering for 10 h at 1600 °C in Ar:H₂ with an addition of Ar:O₂ (about -420 kJ mol⁻¹ oxygen potential) [15]. SEM analyses of the sintered products were carried out as well, and images adapted from those published in the previous study [15] are presented in figure 11. We can see that the sintering at higher temperature for 10 h in the described atmosphere led to intact and spherically shaped particles with no notable cracks for the un-doped and doped particles prepared with the trivalent dopant precursors (figure 11a to 11c). Significantly slower heating rates of 1.5 °C min⁻¹ and 5 °C min⁻¹ were applied during the calcination and sintering, respectively. This might have led to less crack formation during the calcination and the longer sintering time to healing. The Ce-doped particle prepared with Ce^{IV} as dopant precursor (figure 11d) has a deformed spherical shape, but is an integer particle. However, it exhibits a large number of cracks on the surface.

5. Conclusions

The structural changes of un-doped, 10 mol% Nd- and 10 mol% Ce-doped ammonium diuranate microspheres during the conversion into U_{1-y}Ln_yO_{2±x} were investigated in detail by a combination of *in-situ* TG-DSC and HT-SEM. A thermal treatment consisting of a calcination in air at 900 °C for one hour and a sintering under reducing conditions with temperatures up to about 1100 °C for two hours was carried out. Within the heating of the calcination with 10 °C min⁻¹, the formation of cracks and a shrinkage of the particles was observed, slower heating rates (e. g. 1.5 °C min⁻¹) will most likely result in less severe crack formation. The Ce-doped particle prepared with the tetravalent precursor behaved considerably different than the one prepared with the trivalent precursor, pointing to a different crystallinity and/or initial composition and revealed a significant impact of the dopant precursor on the gel formed during the IG synthesis. The shrinkage for higher temperatures is caused mainly by the densification of the material, and the conversion of the UO₃ matrix to an α-U₃O₈ matrix has no noticeable influence. The calcined products were identified as (Ln-doped) α-U₃O₈ by XRD. Upon the introduction of a reducing atmosphere, a sudden mass loss was recognised in the material, while a shrinkage occurred significantly later, pointing out that the shrinkage observed within the sintering is also controlled by a general densification of the material and insignificantly by the volume change due to the transition of the orthorhombic α-U₃O₈ phase to the cubic UO_{2±x} phase. Comparable particle radii during the isotherm at about 1100 °C indicate that the temperature has a higher impact on the shrinkage than the dwell time at a certain temperature. The Nd- and Ce-doped particles prepared with trivalent dopant precursors showed a comparable decomposition and sintering behaviour besides some minor differences during the release of water and ammonia within the calcination. The investigations carried out in this work describe the structural changes of Nd- and Ce-doped ADU microspheres during the conversion to U_{1-y}Ln_yO_{2±x} very detailed for temperatures up to about 1100 °C. When applying commonly used sintering conditions around 1600 °C with longer dwell times, significantly larger volume shrinkage factors can be obtained, leading

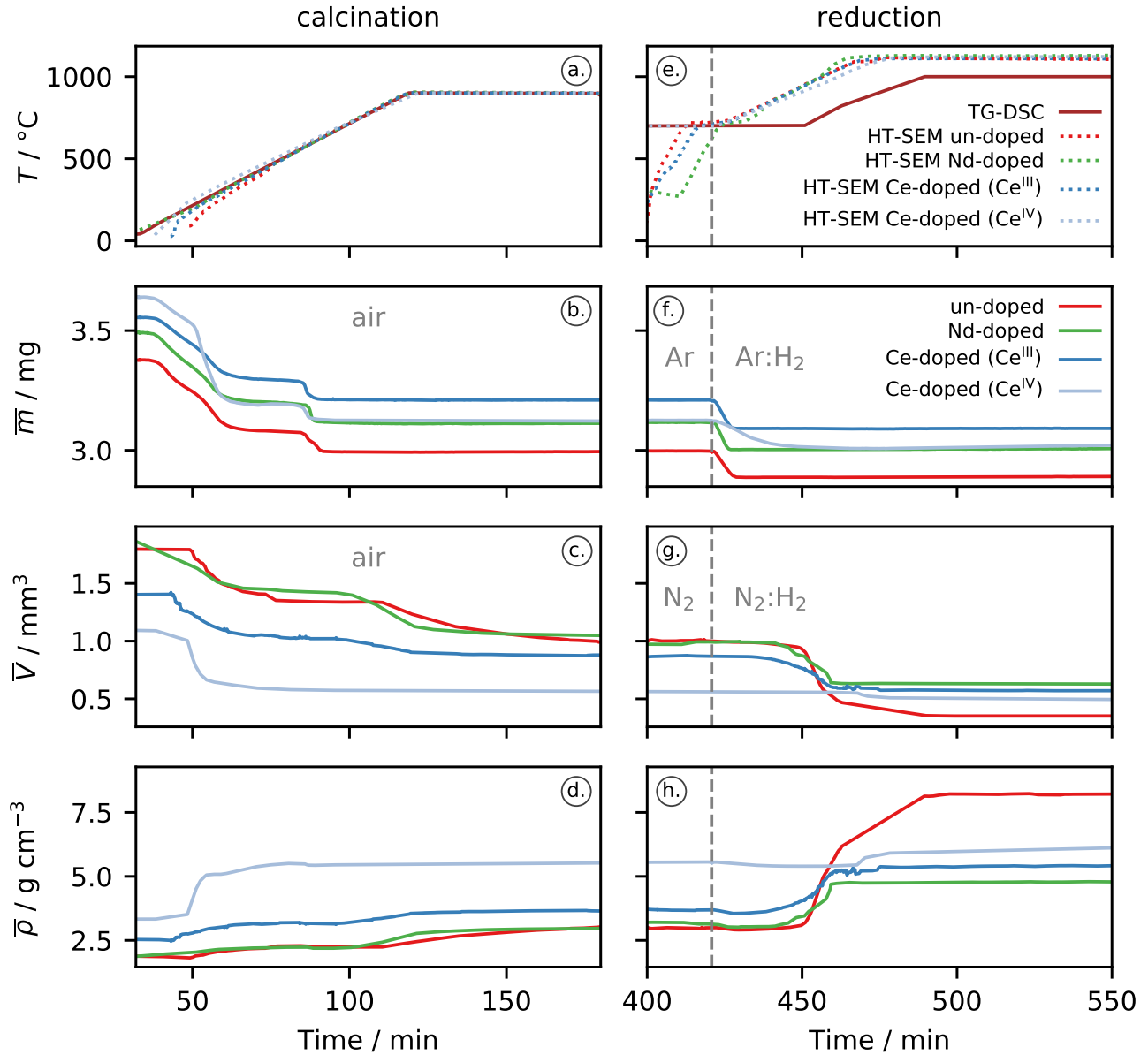


Figure 10: Temperatures measured during the TG-DSC and HT-SEM analyses for the calcination in air (a) and sintering in reducing atmosphere (e). Masses obtained during the TG-DSC measurements standardised to the average particle masses (b, f) and volumes derived from radii measured in the HT-SEM investigations standardised to the average particle volumes (c, g), as well as resulting average *in-situ* particle densities during both thermal treatment steps (d, h).

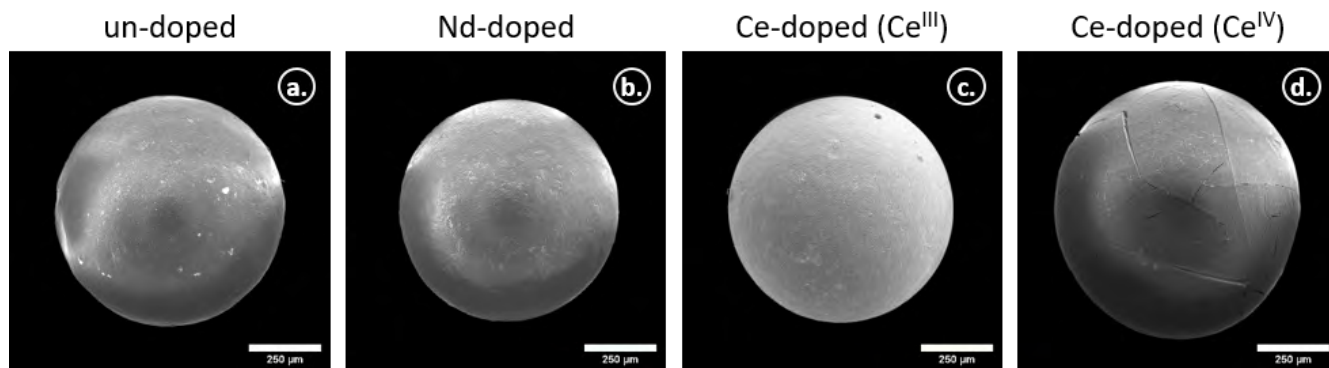


Figure 11: SEM micrographs of the un-doped and 10 mol% Nd- and 10 mol% Ce-doped particles after a sintering for 10 h at 1600 °C in Ar:H₂ (95:5) with an oxygen potential of about 420 kJ mol⁻¹. Images adapted from Schreinemachers et al. [15].

500 to spherical, ceramic particles without cracks on the surface for both, UO_{2±x} microspheres and U_{1-y}Ln_yO_{2±x} microspheres prepared from trivalent dopant precursors.

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Data Availability Statement

The data required to reproduce the findings of this study are available to download from the repository *zenodo* [30, 31].

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