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ARTICLE

Hofmeister Effect in the Keggin-type Polyoxotungstate Series

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The chaotropic character of Keggin-type polyoxotungstate anions was evaluated with respect to their ability to bind to γ -cyclodextrin (γ -CD) by varying the global charge density of the nanometer-sized polyanion. The strengths of the host-guest association were analyzed within the series of isostructural $[XW_{12}O_{40}]^{n-}$ anions where ionic charge varies from 6- to 3- depending on the heteroatom, respectively, $X = H_2^{2+}, B^{3+}, Si^{4+}$ or P^{5+} . Titration experiments using complementary techniques (ITC, DOSY NMR, and electrochemistry) revealed that the affinity between γ -CD and the polyoxometalate (POM) is directly correlated to the charge density of the Keggin anion as reflected in the values of the binding constants $K_{1:1}$. These constants increase dramatically following the order: $[H_2W_{12}O_{40}]^{6-} < [BW_{12}O_{40}]^{5-} < [SiW_{12}O_{40}]^{4-} < [PW_{12}O_{40}]^{3-}$. Additionally, cloud point experiments on a non-ionic surfactant resulted in the same series of POMs emphasize the general affinity of these inorganic Keggin-ions to non-ionic organic soft matter due to a general solvent effect arising from the weakening of the hydration sphere with decreasing ionic charge. Furthermore, single crystal X-ray diffraction analysis showed distinct organizations of the POMs with γ -CD in the solid-state, where the moderate chaotrope $[BW_{12}O_{40}]^{5-}$ interacts with the external wall of the γ -CD, while on the other side of the series, $[PW_{12}O_{40}]^{3-}$ penetrates deeply into the cavity of the γ -CD through its secondary rim offering optimal contact area. Finally, our investigations revealed the unique behavior of $[PW_{12}O_{40}]^{3-}$, which displayed not only the highest affinity of the POMs to γ -CD ($K_{1:1} > 10^5 M^{-1}$), but also the ability to interact with both CD faces resulting in a wide variety of supramolecular aggregates.

Introduction

Polyoxometalates (POMs) are a remarkable class of anionic polynuclear metal oxide clusters of transition metal ions like W^VI or Mo^VI .¹ POMs are able to exchange reversibly and massively electrons without structural change, making them promising electro-active subcomponents of hybrid systems relevant for applications in the energy field.² Recent studies highlighted that POMs are an appealing molecular 'electron-reservoir' to construct lithium-, sodium-ion and redox flow batteries,³⁻⁷ supercapacitors,⁸ fuel cells^{9,10} or catalytic systems¹¹⁻¹⁴ relevant for sustainable processes. The development of straightforward preparation routes to hybrid materials including POM units through supramolecular contacts has led to unique molecular assemblies with novel properties and reactivity.¹⁵ In context, exploration of ionic and other non-covalent interactions (hydrogen bonding, van der Waals forces, etc.) between POMs and other types of

molecular systems has attracted considerable interest, revealing the remarkable supramolecular behavior of POMs to self-assemble across multiple length scales ranging from discrete systems to gigantic spherical single-layered nanostructures known as "blackberries".^{12,16-20}

One of the most striking supramolecular properties of POMs arises from their ability to form in solution resilient non-bonding contacts with neutral surfaces such as organic macrocycles, micelles, proteins or polymers.²¹⁻³⁰ Actually, their solution behavior appears "counter-intuitive" because POMs are anionic inorganic discrete species considered generally as hydrophilic. Others and us established that association phenomena between POMs and organic moieties are driven by a solvent mediated effect arising from a water structure recovery upon release of the POM's and the moiety's hydration water.^{22,26,31} POMs as well as other polynuclear species (dodecaborates or electron-rich octahedral clusters) have been identified as water structure breakers, also named chaotropic entities after the Hofmeister classification.^{22,32-34} Based on the current understanding of the Hofmeister series, the chaotropicity of an ion should increase when its charge density decreases, but regarding POMs more insight is needed into the influence of the ionic charge, especially the effect of redox change on their chaotropic character.^{22,31} From a thermodynamic point of view, processes involving chaotropes are usually considered to be enthalpically driven accompanied by an entropic penalty, i.e. enthalpy-entropy compensation.^{26,35}

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Cyclodextrins (CDs) are natural cyclic oligosaccharides with a hydrophobic internal cavity and a hydrophilic external surface, and thus exhibit interesting properties as host molecules for a wide variety of inorganic species including polyoxometalates.³⁶ CDs are readily available as α -, β - and γ -forms comprising six, seven, or eight glucose units, respectively. The host-guest assembly of the toroidal CDs and spherical POM clusters has attracted much attention and is a fast-growing topic. In 2015, Stoddart and coworkers published the first examples of host-guest complexes between γ - and β -CDs and phosphomolybdate $[\text{PMo}_{12}\text{O}_{40}]^{3-}$.³⁰ Since then, several hybrid (organic/inorganic) adducts or materials have also been reported with various archetypal polyoxotungstates including the Lindqvist-type $[\text{W}_6\text{O}_{19}]^{2-}$,^{37,38} the Keggin-type ions $[\text{PW}_{12}\text{O}_{40}]^{3-}$ and $[\text{SiW}_{12}\text{O}_{40}]^{4-}$,³⁸⁻⁴³ and the Dawson-type $[\text{P}_2\text{W}_{18}\text{O}_{62}]^{6-}$.^{25,26} Some of these hybrid CD/POM materials are promising for application in catalysis^{39,41} and energy storage,⁴⁴ where surface effects and hybrid structure would play crucial role. Therefore, detailed characterization of the interactions between CDs and POMs and a better understanding of the phenomena occurring on their interfaces are essential for further innovative developments of redox-responsive systems and related applications.

and electrochemistry), we show in this article that affinity between γ -CD and Keggin-type anions increases dramatically with decreasing ionic charge. Simultaneously, we evaluate the chaotropic behavior using a simple experimental procedure based on the propensity of the POM species to adsorb on the non-ionic ethylene glycol-based surfactant C_8E_4 . Thus, we establish unambiguously in this report the direct correlation between the chaotropic nature and the formation of highly stable supramolecular complexes. The present study might open new opportunities for the development of smart supramolecular devices with responsive behavior based on redox changing of the POM unit.

Results and discussion

Synthesis

Attempts for isolation of the supramolecular adducts have been carried out systematically using different ratios of sodium salt of POMs $\text{Na}_m[\text{XW}_{12}\text{O}_{40}] \cdot n\text{H}_2\text{O}$ with $\text{X} = \text{P}^{5+}$, Si^{4+} , B^{3+} , H_2^{2+} and varying procedures of crystallization such as slow evaporation of aqueous solution or ethanol vapor diffusion into aqueous solution containing cesium chloride. In all cases, crystalline products were obtained but only $[\text{BW}_{12}\text{O}_{40}]^{5-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$ led to single crystals suitable for XRD analysis. Their composition was established combining different complementary techniques (elemental analysis, EDS, XRD, and TGA). The compounds $\text{CsK}_2\text{H}_2\{[\text{BW}_{12}\text{O}_{40}] \cdot 2(\gamma\text{-CD})\} \cdot 29\text{H}_2\text{O}$ (**BW₁₂•2CD**) and $\text{Na}_3\{[\text{PW}_{12}\text{O}_{40}]@(\gamma\text{-CD})\} \cdot 10\text{H}_2\text{O}$ (**PW₁₂@CD**) present different molecular organization as further discussed in the next section.

Review of Keggin/ γ -CD interactions in the solid-state

The crystal structure of **PW₁₂@CD** was found to be isostructural to the recently published compounds $\text{NaAH}_6\{[\text{PW}_{12}\text{O}_{40}]_3(\gamma\text{-CD})_3\} \cdot n\text{H}_2\text{O}$ ($\text{A} = \text{Co}^{\text{II}}$ or Cu^{II}).⁴⁵ Zhan and coworkers introduced $\text{Co}(\text{NO}_3)_2$ or CuSO_4 to aqueous solutions containing γ -CD and $\text{Na}_3[\text{PW}_{12}\text{O}_{40}]$ to obtain crystalline products.⁴³ In our case, only Na^+ and H^+ were present evidencing that the POM-CD packing in **PW₁₂@CD** is highly flexible to accommodate various types of counter-cations without structural change. **PW₁₂@CD**'s structure consists of a supramolecular 1:1 adduct involving a $[\text{PW}_{12}\text{O}_{40}]^{3-}$ unit partially incorporated into the cavity of γ -CD through its secondary rim (Fig. 2a). This configuration is quite unique among the few known polyoxometalate/ γ -CD composites that mostly feature interactions with the primary face.^{26,30} Of particular interest, comparison with the isostructural Keggin phosphomolybdate $[\text{PMo}_{12}\text{O}_{40}]^{3-}$ reveals a different complexation mode with γ -CD involving the opposite face, namely, the primary rim (Fig. 2b).³⁰ Furthermore, the stoichiometry of the $[\text{PMo}_{12}\text{O}_{40}]^{3-}$:CD adduct is 1:2, while a 1:1 stoichiometry was found with $[\text{PW}_{12}\text{O}_{40}]^{3-}$ although higher stoichiometry numbers should be expected as observed with Dawson-type POM with stoichiometry of 1:1 up to 1:3.²⁶ Indeed, while **PW₁₂@CD** crystallized in a 1:1 solution mixture, other crystalline products were obtained with higher γ -CD contents (2 γ -CDs per POM

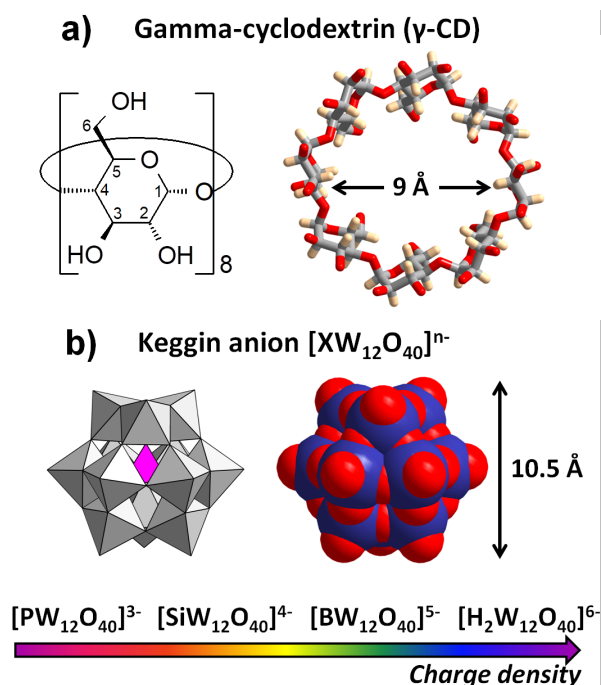


Fig. 1 Structural representations of the molecular entities used as building blocks. (a) γ -cyclodextrin $\text{C}_{48}\text{H}_{80}\text{O}_{40}$ (γ -CD) resulting from the condensation of eight glycopyranose units. The toroidal macrocycle delimits a hydrophobic cavity of 9 Å in diameter. (b) The Keggin-type POM, $[\text{XW}_{12}\text{O}_{40}]^{n-}$, composed of twelve octahedral WO_6 polyhedra that form an anionic tetrahedral cage $\{\text{W}_{12}\text{O}_{40}\}^{8-}$ with a central cavity occupied by the heteroatom X (P^{5+} , Si^{4+} , B^{3+} or H_2^{2+}). The choice of X allows the control of the charge density of the nano-sized inorganic POM.

Herein, we study the formation of host-guest inclusion complexes of γ -CD and Keggin-type POM anions along the $[\text{XW}_{12}\text{O}_{40}]^{n-}$ series with anionic charges from 3- to 6- (Fig. 1) while keeping constant the shape and size of the host and of the guest. Using a set of complementary techniques (NMR, ITC

unit), but we were unable, despite numerous repeated attempts, to solve the solid-state structures of these crystals featuring a tetragonal unit cell ($a = b = 23.94$ and $c = 18.54$ Å). The 1:1 adduct observed in **PW₁₂@CD** exhibits several non-bonding contacts involving the terminal O^t and bridging O^b atoms of the POM and H3 and H5 protons (see Fig. 1a for H labeling) of the inner cavity of the γ -CD (see Fig. S1, in ESI). The short interatomic distances H3...O^t=W (from 2.69 to 3.31 Å), H3...O^b-W (from 2.41 to 2.88 Å), H5...O^t=W (from 2.29 to 3.38 Å), and H5...O^b-W (from 3.13 to 3.88 Å) reflect a strong and deep host-guest inclusion. As **PW₁₂@CD** crystallizes in a cubic crystallographic system, the 1:1 host-guest adduct is surrounded by four host-guest units giving additional non-bonding contacts between the exposed face of the Keggin ion and the external surface of each of the neighboring γ -CDs (see Fig. S2, in ESI).

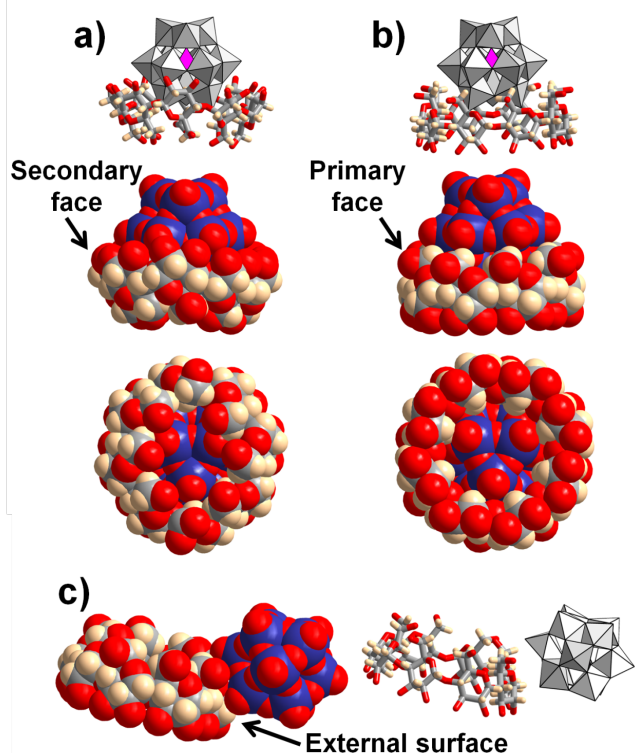


Fig. 2 Illustration of the different supramolecular assemblies observed in the solid-state involving γ -CD and the Keggin anions (a) [PW₁₂O₄₀]³⁻ in **PW₁₂@CD**, (b) [PMo₁₂O₄₀]³⁻ from reference ³⁰, and (c) [BW₁₂O₄₀]⁵⁻ in **BW₁₂*2CD**.

In the solid-state, the organization of the [BW₁₂O₄₀]⁵⁻ and γ -CD is significantly different since the crystalline structure of **BW₁₂*2CD** reveals that the two components co-crystallize without forming an inclusion complex. Indeed, the [BW₁₂O₄₀]⁵⁻ is positioned near the γ -CD in close contact with its external wall (Fig. 2c). One [BW₁₂O₄₀]⁵⁻ unit is surrounded by eight γ -CDs and each γ -CD is in contact with four [BW₁₂O₄₀]⁵⁻ through the three outward protons H1, H2, and H4 as well as the H6 of the methoxy arm (see Fig. S3, ESI). This leads to a global stoichiometry of 2(γ -CD):1POM. Two neighboring γ -CDs are face-to-face connected through a hydrogen bonding network from their secondary rims. The [BW₁₂O₄₀]⁵⁻ is therefore

positioned near the primary face with H6...O^t=W and H6...O^b-W distances in the 2.66-2.91 Å and 2.73-3.91 Å range, respectively. This arrangement with empty γ -CDs is unique in POM/ γ -CD composites although it could occur with the smaller α -CD.⁴⁴ Indeed, the [PW₁₂O₄₀]³⁻ Keggin anion led to an extended framework with α -CD forming double layers of interconnected [PW₁₂O₄₀]³⁻ and CD motifs through bridging alkali cations.⁴⁴ This highlights the importance of the size-matching effect upon the molecular recognition process,⁴⁶ however this could not explain the difference of binding interactions observed in Keggin/ γ -CD arrangement.

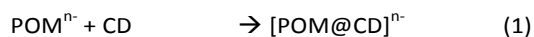
Finally, it is surprising to see how γ -CD interacts so differently with isostructural nano-sized objects, namely [PW₁₂O₄₀]³⁻, [PMo₁₂O₄₀]³⁻, and [BW₁₂O₄₀]⁵⁻. By keeping the charge constant, the basicity of Keggin molybdates is known to be stronger compared to their tungstate analogues as the charge density is higher on the oxygen atoms at the surface of the molybdates.^{47,48} Furthermore, the ionic charge must also play a crucial role in the supramolecular association.²¹ Both these aspects are expected to affect significantly solvent structuration around the POM, and should have important consequences in the formation and stability of the resulting POM-based supramolecular arrangement. Systematic studies in solution are thus needed to shed light on such effects.

Stability in solution

To investigate the complexation of the Keggin-type anions by γ -CD, titration experiments were conducted in solution by means of a set of complementary techniques including NMR spectroscopy (¹H 1D and DOSY), electrochemistry, and isothermal titration calorimetry (ITC).

Isothermal titration calorimetry

ITC experiments were carried out to determine the thermodynamical fingerprint of the interaction of γ -CD respectively with the four Keggin-type heteropolytungstates. ITC thermograms and isotherms are shown in the Electronic Supplementary Information (see Figs. S4-6, ESI), and a summary of binding constants, enthalpy $\Delta_r H^*$ and entropy $\Delta_r S^*$ changes are given at 298 K in Table 1. Beside the 1:1 complexation, a second 1:2 step process is systematically considered, according to equations (1) and (2) shown below.



Nonetheless, only ITC data for silicotungstate were found consistent with the 2:1 binding model involving a sequential process. No significant heat exchange was observed in the case of [H₂W₁₂O₄₀]⁶⁻ indicating negligible interactions with γ -CD. For all other complexes, supramolecular binding revealed to be enthalpically driven. The enthalpic gain is accompanied by an entropic penalty according to the usual enthalpy-entropy compensation observed within CD-based host-guest series.⁴⁹ These thermochemical fingerprints together with low values of both enthalpic and entropic changes suggests that Keggin

anions behave as chaotropic species (water structure breakers).³¹ In such a case, the magnitude of the chaotropic character and the related binding constants are expected to increase with decreasing global charge. However, the highest enthalpic and entropic changes (and binding constants), were observed for $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and the lowest values were found for $[\text{PW}_{12}\text{O}_{40}]^{3-}$. This result represents therefore an anomaly within the expected series, as confirmed by the electrochemical and NMR studies discussed below. The discrepancy here may have appeared for two reasons: i) the partial hydrolysis of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ anion in the ITC conditions (non-buffered and diluted solution) or ii) the simultaneous formation of numerous complexes with various stoichiometries, e.g. 1:1, 1:2, 2:1, etc., (overparametrization).

Table 1 Binding constants (M^{-1}) of γ -CD with the Keggin-type POMs and associated thermodynamic parameters ($\text{kJ}\cdot\text{mol}^{-1}$) at $T = 298\text{ K}$ measured by ITC. Calculated uncertainties are inferior to 10%.

POM	POM:CD	K	$\Delta_r H^*$	$T\Delta_r S^*$	$\Delta_r G^*$
$[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$	-	-	-	-	-
$[\text{BW}_{12}\text{O}_{40}]^{5-}$	1:1	1032	-53.5	-36.3	-17.2
$[\text{SiW}_{12}\text{O}_{40}]^{4-}$	1:1	17209	-56.4	-32.3	-24.2
	1:2	459	-55.0	-39.8	-15.2
$[\text{PW}_{12}\text{O}_{40}]^{3-}$	1:1	2919	-35.1	-15.3	-19.8

Electrochemical studies

Interactions of the Keggin-type ions with γ -CD have been studied further using cyclic voltammetry. As shown below, the presence of γ -CD alters significantly the redox properties of the POMs. Besides, the POM's redox state is expected to have high impact on the binding constants. As shown in Fig. 3, dependence of CVs with γ -CD appears strongly dependent on the nature of the Keggin anion.

For instance, as the amount of γ -CD increases up to 8 equivalents, the half-wave potential related to the two first monoelectronic transfers of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion decreases continuously (see Fig. 3). Actually, the electrochemical behavior of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion gives the most representative effect reflected by a rough dependency of about $-60\text{ mV}/p\text{CD}$ where $p\text{CD} = -\log[\text{CD}]$ (see Fig. S9, ESI). This variation indicates that γ -CD is involved in the predominant redox process written in equation (3), with two consecutive monoelectronic transfers, $n = 3$ or 4 for the first and the second redox processes, respectively. Furthermore, the complexation behavior of γ -CD should be rather consistent with $x = 2$ or 1 corresponding to a 2:1 and 1:1 supramolecular adduct, respectively.

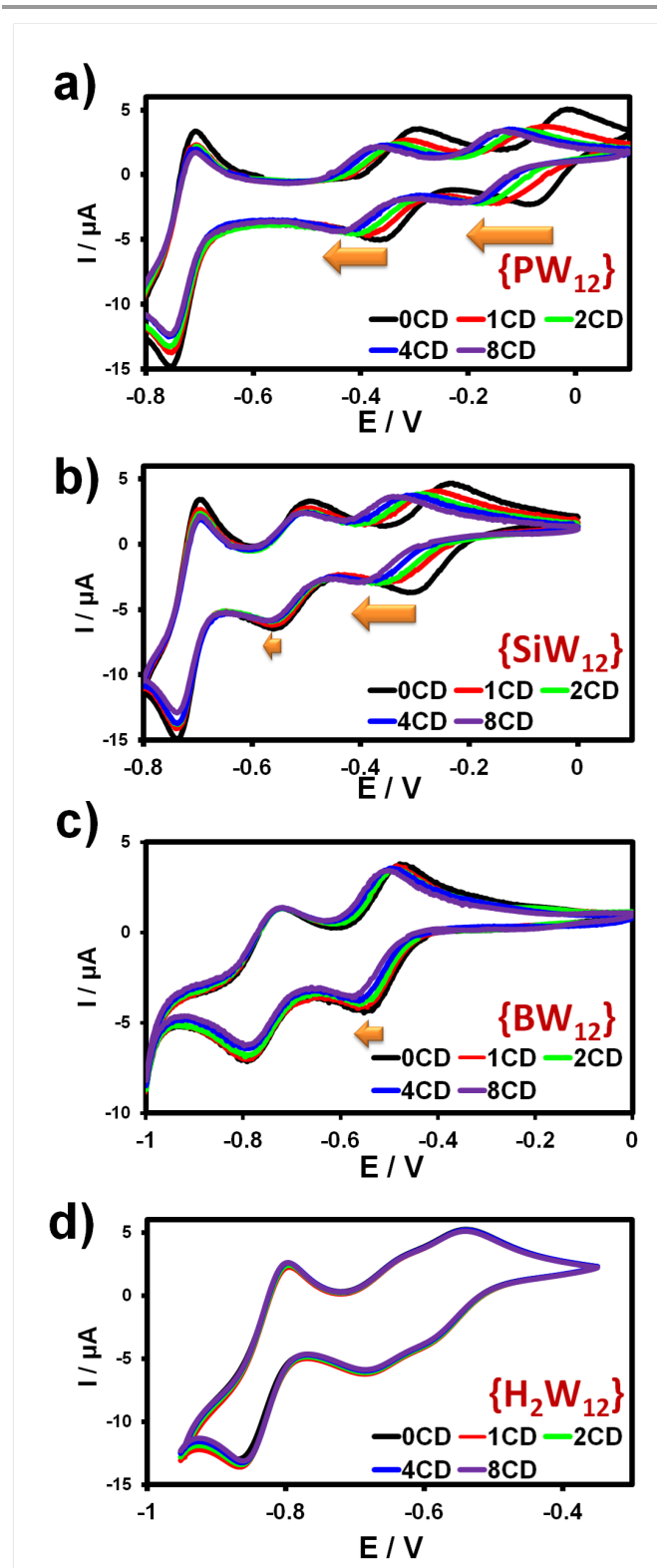
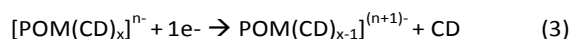
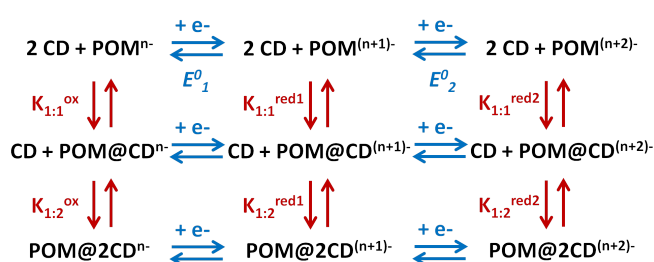


Fig. 3 Cyclic voltammetry of the $[\text{XW}_{12}\text{O}_{40}]^{n-}$ anions with (a) $\text{X} = \text{P}^{5+}$, (b) Si^{4+} , (c) B^{3+} or (d) H_2^{2+} ($0.5\text{ mmol}\cdot\text{L}^{-1}$, glassy carbon working electrode, scan rate $50\text{ mV}\cdot\text{s}^{-1}$) in the presence of increasing amounts of γ -CD (from 0 to 8 equivalents). The experiments involving $[\text{PW}_{12}\text{O}_{40}]^{3-}$, $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and $[\text{BW}_{12}\text{O}_{40}]^{5-}$ have been performed in $25\text{ mmol}\cdot\text{L}^{-1}$ HClO_4 aqueous solution, while acetate buffer solution (0.5 M CH_3COONa ; 0.5 M CH_3COOH) has been used for the electrochemical measurements of $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$.

Along the Keggin series, the variation of the redox properties becomes less pronounced as the ionic charge of the Keggin ion increases. Thus, in the case of $[\text{SiW}_{12}\text{O}_{40}]^{4-}$, the first mono-electronic redox wave appears strongly affected, whereas the second one varies weakly. For the boron derivative, only the first redox wave is moderately shifted and finally, the redox features of $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ remain nearly unchanged in the presence of γ -CD (see Fig. 3). With the exception of $[\text{PW}_{12}\text{O}_{40}]^{3-}$, the redox behavior reflects the trend highlighted by the ITC analysis.

Quantitative analysis of these electrochemical data can yield the equilibrium constants ($K_{1:1}$ and $K_{1:2}$) regarding the binding of the Keggin anions at various redox states by the γ -CD. Considering the three available oxidation states of POM^{n-} (ox), $\text{POM}^{(n+1)-}$ (red₁), $\text{POM}^{(n+2)-}$ (red₂), associated to the three related complexed forms $\text{POM}@x\text{CD}$ with $x = 0, 1$ or 2 , the 3X3 matrix diagram can be drawn (see scheme 1).



Scheme 1 Equilibria involved in redox-dependent supramolecular association of γ -cyclodextrin (CD) and the Keggin type anions (POM).

Including the six redox components involved in each redox process, the Nernst equations are written as followed:

$$E_{\text{peak},1} = E_1^0 + \frac{RT}{F} \ln \frac{(1 + K_{1:1}^{\text{ox}} [\text{CD}]_{\text{eq}} + K_{1:1}^{\text{ox}} K_{1:2}^{\text{ox}} [\text{CD}]_{\text{eq}}^2)}{(1 + K_{1:1}^{\text{red1}} [\text{CD}]_{\text{eq}} + K_{1:1}^{\text{red1}} K_{1:2}^{\text{red1}} [\text{CD}]_{\text{eq}}^2)} \quad (4)$$

$$E_{\text{peak},2} = E_2^0 + \frac{RT}{F} \ln \frac{(1 + K_{1:1}^{\text{red1}} [\text{CD}]_{\text{eq}} + K_{1:1}^{\text{red1}} K_{1:2}^{\text{red1}} [\text{CD}]_{\text{eq}}^2)}{(1 + K_{1:1}^{\text{red2}} [\text{CD}]_{\text{eq}} + K_{1:1}^{\text{red2}} K_{1:2}^{\text{red2}} [\text{CD}]_{\text{eq}}^2)} \quad (5)$$

where E_1^0 or E_2^0 are the standard potentials of each redox process and $[\text{CD}]_{\text{eq}}$, the concentration in free γ -CD at the equilibrium. The values $K_{1:1}$ and $K_{1:2}$ correspond to the binding constants associated to equation (1) and (2), respectively, wherein the POM can be in its three redox states, oxidized (ox), one-electron reduced (red₁) or two-electron reduced (red₂) (see scheme 1). Depending on the Keggin-type POM, equations (4) or (5) were adapted and then used to extract the binding constants through the fitting of experimental shift of the apparent potential upon free γ -CD concentration (see ESI, Figs. S7-9). The case of the metatungstate ion $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ is trivial as neither ITC nor electrochemistry evidences any interaction with γ -CD. For the $[\text{BW}_{12}\text{O}_{40}]^{5-}$ ion, ITC experiments showed formation of a 1:1 adduct. Furthermore, the second wave is nearly independent of the γ -CD concentration meaning that the one-electron and the two-electron reduced forms do not interact with γ -CD. In such conditions, the expression of the apparent potential $E_{\text{peak},1}$ is adapted from equation (4) by using only the binding constant $K_{1:1}^{\text{ox}}$. Decreasing further the ionic charge produces a richer and more complex behavior.

The first wave of the $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ anion undergoes a shift of about -60 mV per $p\text{CD}$ unit (see ESI, Fig. S8) meaning the Nernst Equation can be derived from equation (3). The second wave exhibits only moderate variation, consistent with a weak complexation of the one-electron reduced form (red₁) and no complexation for the two-electron reduced form (red₂). Then the equation (5) is derived by using only the binding constant $K_{1:1}^{\text{red1}}$ and matches the experimental data for $K_{1:1}^{\text{red1}} = 130 \text{ M}^{-1}$. Considering the first monoelectronic exchange, both the 1:1 and the 1:2 adduct must be considered for the oxidized form as shown by ITC. Then equation (4) should include $K_{1:1}^{\text{ox}}$, $K_{1:2}^{\text{ox}}$ and $K_{1:1}^{\text{red1}} = 130 \text{ M}^{-1}$. Suitable fitting with experimental data was obtained with numerical values 9500 and 150 M^{-1} for $K_{1:1}^{\text{ox}}$ and $K_{1:2}^{\text{ox}}$, respectively. Finally, the electrochemical behavior of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion consists of two CD-dependent monoelectronic waves which move both by about -60 mV per $p\text{CD}$ unit in accordance with equation (3). The second wave is satisfactory modeled by using $K_{1:1}^{\text{red1}} = 6000 \text{ M}^{-1}$ and $K_{1:2}^{\text{red1}} = 30 \text{ M}^{-1}$ while calculation of the first redox potentials requires to include $K_{1:1}^{\text{ox}}$ and $K_{1:2}^{\text{ox}}$ in addition to the binding constants related to the reduced form $[\text{PW}_{12}\text{O}_{40}]^{4-}$, namely red₁. From equation (4), good agreement was found between calculated and experimental data for $K_{1:1}^{\text{ox}} = 90000 \text{ M}^{-1}$ and $K_{1:2}^{\text{ox}} = 1500 \text{ M}^{-1}$. The overall binding constants values are reported in Table 2 and the comparison between experimental data and the simulated evolution of the observed apparent potential is reported in ESI (see Figs. S7-9). The binding constants determined by electrochemistry for $[\text{BW}_{12}\text{O}_{40}]^{5-}$ and $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ ions are significantly lower compared to those found by ITC (see Table 1). This difference is likely due to the different medium used in the ITC experiments (pure water) and in the electrochemical measurements (aqueous solution of $0.025 \text{ mol.L}^{-1} \text{ HClO}_4$) and may arise from the presence of hydronium ions H_3O^+ or perchlorate ion ClO_4^- involved in competition processes with the POM-CD aggregation. Importantly, analysis of the electrochemical data allows us to determine the $K_{1:1}$ binding constants of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion. With a value of $9 \times 10^4 \text{ M}^{-1}$, the affinity of $[\text{PW}_{12}\text{O}_{40}]^{3-}$ for γ -CD is about one order of magnitude higher than the affinity of $[\text{SiW}_{12}\text{O}_{40}]^{4-}$. Such a high value corresponds to one of the largest affinity constants reported so far with γ -CD.^{32,50,51} Moreover, the second binding constant $K_{1:2}$ is also quite large, around one order of magnitude higher than the $K_{1:2}$ value observed for $[\text{SiW}_{12}\text{O}_{40}]^{4-}$.

Table 2 Binding constants $K_{1:1}$ and $K_{1:2}$ involving γ -CD with the Keggin-type anions at different oxidation states as determined from electrochemistry. The accuracy is estimated to be within 10% of the corresponding value.

X	POM:CD	$[\text{XW}_{12}\text{O}_{40}]^{3-}$	$[\text{XW}_{12}\text{O}_{40}]^{4-}$	$[\text{XW}_{12}\text{O}_{40}]^{5-}$
P	1:1	90 000	6 000	0
	1:2	1 500	30	0
Si	1:1	-	9 500	130
	1:2	-	150	0
B	1:1	-	-	600
	1:2	-	-	0

This electrochemical study enabled to restore a logical sequence within the Keggin-type series, corroborating the fact that the POM/ γ -CD association is mainly related to the global charge of POMs. For instance, plotting $\text{Log}K_{1:1}$ versus the ionic charge of the Keggin-type anion results in a quite perfect linear relationship (see ESI, Fig. S10). Furthermore, analysis of the electrochemical data provides access to the binding constants of the lower oxidation states of the Keggin anions (see Table 2). The binding constants of the reduced POM⁽ⁿ⁺¹⁾⁻ exhibit lower values compared to those determined for the oxidized parent (POMⁿ⁻), evidencing again the preponderant effect of the charge on the association process with the macrocyclic host. For instance, the binding constants of the one-electron reduced $[\text{PW}_{12}\text{O}_{40}]^{4-}$ anion ($K_{1:1} = 6000 \text{ M}^{-1}$) fall in similar range order to that found for the oxidized silicotungstate $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ ($K_{1:1} = 9500 \text{ M}^{-1}$) (see Table 2). A similar observation is made for the one electron-reduced silicotungstate $[\text{SiW}_{12}\text{O}_{40}]^{5-}$ and the oxidized borotungstate $[\text{BW}_{12}\text{O}_{40}]^{5-}$, which both exhibit comparable affinity toward γ -cyclodextrin. Besides, such a trend can be compared to results reported in a previous study which demonstrated that decreasing the ionic charge of a rhenium-selenide cluster $[\text{Re}_6\text{Se}_8(\text{CN})_6]^{4-}$ through a one-electron oxidative process provokes a dramatic increase of the $K_{1:1}$ binding constant from 1.5×10^3 to $2.2 \times 10^5 \text{ M}^{-1}$.³⁴ Nevertheless, alteration of the redox properties through supramolecular interactions remains quite intriguing in the case of the Keggin-type polyoxometalates. γ -CD being a non-ionic component, its interaction with POM does not change the ionic charge as observed generally with protons or lithium ions. In context, Proton-Coupled Electron Transfer (PCET) involving POM species is a well-known pH dependent process featured by an increase of the standard potential as pH decreases. Such an effect can even be extended to POM-lithium aggregation process under reducing conditions in non-aqueous solvents.⁵² In the presence of γ -CD, the origin of the redox potential variation upon γ -CD interaction should arise mostly from a change in the solvation shell of the POM corresponding to a shed of surrounding water molecules upon binding to γ -CD resulting in a more hydrophobic environment around the POMs. The influence of the medium, *i.e.* solvent and electrolyte support, on the redox properties of the POM species is well documented experimentally as well as through theoretical calculations and related studies reinforce our interpretation of the γ -CD effect upon the redox properties of Keggin-type POMs.^{53,54} For instance, the first one-electron

wave of the $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ ion undergoes a 600-700 mV potential shift toward the negative potential when the medium is changed from aqueous to organic such as acetonitrile or dimethyl sulfoxide. This effect has been attributed to the capacity of the solvent to stabilize the reduced derivatives and has been quantified by the acceptor number (AN) of the solvent showing that the reduction process becomes difficult as solvent AN decreases. Water exhibits a high AN value (58.4) while less hydrophilic or less polar media feature lower AN parameters.⁵⁵ Furthermore, the intensity of the peak current decreases significantly in the presence of γ -CD. This behavior accounts for the complexation process which lowers the diffusion coefficient of the parent oxidized POM and then consecutively affects the peak current intensity, according to the Randles-Sevcik equation.⁵⁶ This effect on the current peak intensity appears quite pronounced in the CVs of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion, which exhibits the strongest affinity with the γ -CD and diminishes as the ionic charge of the POM increases. Diffusion coefficient change upon γ -CD complexation is also nicely supported by ¹H DOSY NMR study presented below.

To sum up, electrochemical investigations demonstrate that the binding constants $K_{1:1}$ and $K_{1:2}$ increase by at least one order of magnitude as the global ionic charge decreases along the Keggin-type series $[\text{XW}_{12}\text{O}_{40}]^{n-}$. From this observation, the Keggin/ γ -CD association appears as an appealing component for redox switches, and could be integrated into design stimuli-responsive supramolecular systems controlled by the redox state of the POMs.

NMR spectroscopy

The formation and stability of Keggin POM/ γ -CD complexes in aqueous solution were further probed by ¹H NMR spectroscopy in D₂O. In previous separate studies, the interaction of γ -CD with both $[\text{PW}_{12}\text{O}_{40}]^{3-}$ and $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ has been characterized by ¹H NMR.^{39,45} Here, we conducted a comparative titration study of 2 mM aqueous γ -CD solution with the four Keggin-type anions. Fig. 4 shows the resulting ¹H NMR spectra in the range 3.7-4.5 ppm corresponding to the resonance domain of the most exposed protons to guest species, namely H3 and H5 are located inside the cavity and H6 on the top of the secondary rim (see Fig. 1 for protons numbering scheme and ESI Figs. S11-14 for the complete ¹H NMR spectra). These protons are the most sensitive to host-guest contact and are usually used as indicators to probe the inclusion phenomenon.

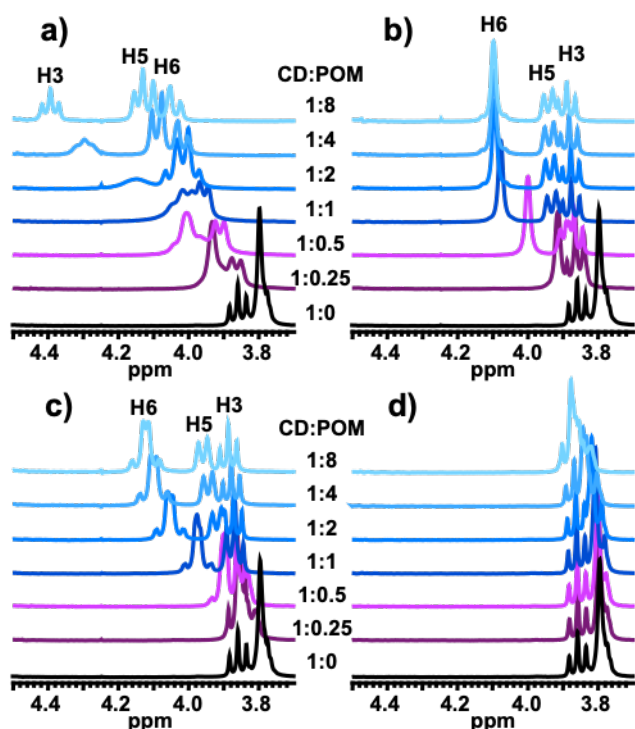


Fig. 4 ^1H NMR spectra in the chemical shift range for H3, H5, and H6 resulting from the titration of 2 mmol.L $^{-1}$ aqueous γ -CD solution with $[\text{XW}_{12}\text{O}_{40}]^n$ POMs. (a) $[\text{PW}_{12}\text{O}_{40}]^{3-}$; (b) $[\text{SiW}_{12}\text{O}_{40}]^{4-}$; (c) $[\text{BW}_{12}\text{O}_{40}]^{5-}$ and (d) $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$.

Depending on the Keggin anion, the spectra underwent more or less strong alteration evidencing POM-specific recognition. Table 3 summarizes the extent of shift variations observed for H3, H5, and H6 after adding eight equivalents of each Keggin-type POM. As expected, the ^1H NMR spectrum of the macrocyclic host remains nearly unchanged when $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ was added to the γ -CD solution, and the maximum shifts observed were ca. 0.07 ppm for both H5 and H6 indicating very weak interaction. Notably, much bigger effects were observed in the cases of $[\text{BW}_{12}\text{O}_{40}]^{5-}$ and $[\text{SiW}_{12}\text{O}_{40}]^{4-}$, where shifts up to 0.33 and 0.18 ppm were observed for H6 and H5, respectively. These changes in the spectra are similar that suggests comparable behavior in solution for $[\text{BW}_{12}\text{O}_{40}]^{5-}$ and $[\text{SiW}_{12}\text{O}_{40}]^{4-}$. Nevertheless, close inspection revealed some differences. First, the H6 (and also H5) signal evolves much more significantly with increasing the $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ amount compared to $[\text{BW}_{12}\text{O}_{40}]^{5-}$ although the final chemical shift at the highest ratio is the same. As the observed chemical shift corresponds to a weighted average of free and complexed γ -CD species, this means the binding constant of POM-CD complexation should be higher with $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ than with $[\text{BW}_{12}\text{O}_{40}]^{5-}$. The second observation is the splitting of the initial singlet of the H6 protons into two unresolved doublets in $[\text{BW}_{12}\text{O}_{40}]^{5-}$ containing solution. This may be due to constrained dynamic rotation of the methoxy arms which narrows the signal of the two methylenic diastereotopic H6 protons. Such steric hindrance would result from the interaction of the POM specifically with the primary rim of the γ -CD since the H3 protons present in the secondary

rim did not show any significant perturbation. The situation is totally different with $[\text{PW}_{12}\text{O}_{40}]^{3-}$ anion where besides the shift of H6, we can also appreciate strong effects on both H5 and H3 resonances with up to 0.53 and 0.36 ppm downfield shifts, respectively, indicating involvement of the secondary face of the γ -CD in the strong interaction with $[\text{PW}_{12}\text{O}_{40}]^{3-}$. Furthermore, we note the line broadening of H3 signal for POM: γ -CD ratio lower than two which is the signature of strong POM/ γ -CD interactions involved in a 1:2 stoichiometry supramolecular adduct.⁵⁷

Table 3 Differences in chemical shifts (ppm) of protons H3, H5, and H6 of γ -CD before and after addition of eight equivalents of POM.

POM	$\Delta\delta\text{H3}$	$\Delta\delta\text{H5}$	$\Delta\delta\text{H6}$
$[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$	0.02	0.06	0.07
$[\text{BW}_{12}\text{O}_{40}]^{5-}$	0.03	0.18	0.33
$[\text{SiW}_{12}\text{O}_{40}]^{4-}$	0.03	0.16	0.30
$[\text{PW}_{12}\text{O}_{40}]^{3-}$	0.53	0.36	0.28

By analyzing the chemical shifts variation arising from fast chemical exchange regime, quantitative treatment allows for a determination of the binding constants for 1:1 POM: γ -CD complexes (for further details, see Figs. S15-16, ESI). The values obtained for $K_{1:1}$ are 23 and 1100 M $^{-1}$ for $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$:CD and $[\text{BW}_{12}\text{O}_{40}]^{5-}$:CD, respectively, that compare well with the ITC and electrochemistry results (see Tables 1 and 2). In summary, NMR results are fully consistent with electrochemistry showing that the interaction strength with γ -CD increases following this order: $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-} < [\text{BW}_{12}\text{O}_{40}]^{5-} < [\text{SiW}_{12}\text{O}_{40}]^{4-} < [\text{PW}_{12}\text{O}_{40}]^{3-}$.

Further quantitative investigations were also carried out using DOSY NMR. This technique is a valuable tool to study dynamic equilibria, and similarly to chemical shift, the observed diffusion coefficient D in the fast exchange limit can be described as a weighted average of a given species evolving between different states.⁵⁸ Thus, the NMR observable of γ -CD in presence of POM corresponds to an average situation weighted on residency times between free (solvated) and bound states. The observed diffusion coefficients D as a function of POM: γ -CD molar ratio produce different profiles depending on the nature of the Keggin-type anions (see Fig. 5). From the initial value for solvated γ -CD measured at 254 $\mu\text{m}^2/\text{s}$, the self-diffusion D drops more or less abruptly upon addition of POM and reaches a plateau at large excess of POM at 226 $\mu\text{m}^2/\text{s}$ in the case of $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and $[\text{BW}_{12}\text{O}_{40}]^{5-}$, while for the $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ ion, the decrease of the self-diffusion coefficient draws a smooth slope featured by larger self-diffusion values as an indication of very weak POM- γ -CD interactions. For the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ions, the D profile is characteristic of strong interactions. The decrease of D is sharper, reaching even the lower limit value measured at 219 $\mu\text{m}^2/\text{s}$. We note also a minimum value for the curves of both $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$ at ca. POM: γ -CD = 0.5 suggesting presence of bulkier intermediate species at these compositions. Both curve profiles have been modeled

satisfactorily using either a one-step complexation model for $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ and $[\text{BW}_{12}\text{O}_{40}]^{5-}$ or a two-step complexation model for $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$. Simulations of these curves considering binding constants consistent with those determined previously from ITC, electrochemistry and also 1D NMR studies allowed to estimate diffusion coefficients of γ -CD within 1:1 and 1:2 adducts. Fair agreements between calculated and experimental data were found using D values of 226 and 163 $\mu\text{m}^2/\text{s}$ for POM@CD and POM@2CD adducts, respectively. Noteworthy, the same D values were used for the 1:1 or 1:2 adduct independently of the nature of the Keggin-type anion (see ESI for details of modeling, Figs. S17-20), considering similar volume/size for the POM-CD adduct along the Keggin-type series. The observed lower limit value for D with $[\text{PW}_{12}\text{O}_{40}]^{3-}$ (219 vs 226 $\mu\text{m}^2/\text{s}$, see ESI for modeling, Fig. S20) highlights the remarkable solution behavior of the phosphotungstate ion as reported by Antonio et al.⁵⁹ In this composition domain, it may result from some further aggregation which could be consistent with a γ -CD unit sandwiched by two $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ions interacting with both primary and secondary faces of the γ -CD. This hypothesis is also quite consistent with the ^1H NMR titration which reveals interactions involving simultaneously both rims of the γ -CD.

Furthermore, ^1H NMR can also characterize the proton-containing $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ anion exhibiting a resonance at ca. 6.0 ppm. Very small changes in the diffusion coefficient D (245–260 $\mu\text{m}^2/\text{s}$) were observed over the titration experiment, reflecting very weak interaction with γ -CD as previously identified.

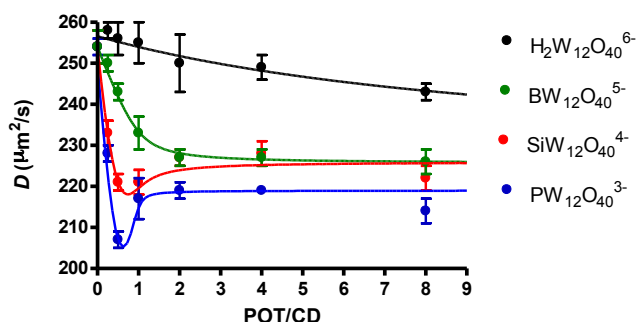


Fig. 5 Diffusion coefficients of γ -CD as a function of POM: γ -CD molar ratio, measured from ^1H DOSY NMR of a solution of 2 mM of γ -CD with various POM content. Lines correspond to calculated curves with one-step 1:1 POM:CD complexation for $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ and $[\text{BW}_{12}\text{O}_{40}]^{5-}$, and with two-step 1:2 POM:CD complexation for $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ and $[\text{PW}_{12}\text{O}_{40}]^{3-}$.

Driving force of the aggregation processes

This comparative study evidences clearly that Keggin-type anions bind to γ -CD with different extents despite their isostructure. Although γ -CD is non-ionic, the charge of the POM plays a dramatic role in the molecular recognition process, both in its thermodynamic stability and in its supramolecular structure. All experimental measurements (ITC, electrochemistry and NMR) have shown that the interaction with $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ is very weak ($K_{1:1} \approx 2.3 \times 10^1 \text{ M}^{-1}$), while it is very strong with $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ($K_{1:1} \approx 1.6 \times 10^5 \text{ M}^{-1}$). It was

proposed that the binding affinity is related to the charge density of the polyanion, which has an effect on the hydration strength, *i.e.* hydration free energy of the ions. The monotonous trend observed along the oxidized derivatives of the Keggin-type series could be understood by thermodynamic features of the hydration shell around the POM, which is directly related to its chaotropic nature. The weak charge density of the Keggin-type POM generates a loosely bound solvation shell of disordered water molecules with high structural entropy. The energy gain associated to the release of this hydration water into the bulk upon binding to non-ionic moieties (at micellar surfaces, water surfactant monolayer or macrocycles) was proposed as the main driving force of the superchaotropic effect of POMs and more generally of low charge density nano-ions.²² The evolution of the cloud point of a non-ionic surfactant, tetra-ethylene glycol mono-octyl ether (C_8E_4), upon addition of POMs was proposed as a tool to evaluate their chaotropic behavior²². The higher the observed cloud point increase, the stronger the interaction with the non-ionic surfactant and the stronger the chaotropic nature of the tested ion. This procedure was carried out for the Keggin-type anions investigated here. Corresponding to the respective increase in the cloud point by each POM, Fig. 6 illustrates an increase in the chaotropic nature of the polyanions from $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ to $[\text{PW}_{12}\text{O}_{40}]^{3-}$, *i.e.* from high to low charge density. The superchaotropic P-, Si- and B-Keggin POMs produce much stronger increases in the cloud point compared to the classical chaotrope SCN^- . The resulting ordering of POMs according to their chaotropic behavior is in full agreement with the series obtained by electrochemistry and NMR measurements.

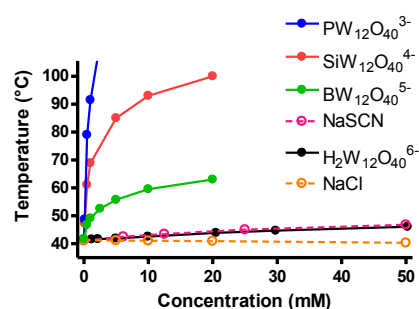


Fig. 6 Cloud point evolutions of 60 mM C_8E_4 as a function of salt concentration for the four isostructural Keggin-ions $\text{PW}_{12}\text{O}_{40}^{3-}$, $\text{SiW}_{12}\text{O}_{40}^{4-}$, $\text{BW}_{12}\text{O}_{40}^{5-}$ and $\text{H}_2\text{W}_{12}\text{O}_{40}^{6-}$. The classical salts NaCl (neutral salt) and NaSCN (classical chaotropic salt) are included as references.

SAXS measurements provided further insight into the superchaotropic character of the Keggin POMs (see Figs. S21 and S22, ESI). $\text{H}_2\text{W}_{12}\text{O}_{40}^{6-}$ in the presence of C_8E_4 shows identical scattering as in bare water, while with $\text{BW}_{12}\text{O}_{40}^{5-}$ a strong oscillation appears, which is even more pronounced with $\text{SiW}_{12}\text{O}_{40}^{4-}$. This oscillation is characteristic of core-shell structuration and typical of POM-decorated micelles. The increase of the oscillation with the superchaotropic character of the Keggin-ion is a direct indicator of the extent of adsorption onto the C_8E_4 -micelles resulting in the same

ordering as in other experiments: $\text{SiW}_{12}\text{O}_{40}^{4-} > \text{BW}_{12}\text{O}_{40}^{5-} > \text{H}_2\text{W}_{12}\text{O}_{40}^{6-}$. It is worthy to note that in the case of $\text{PW}_{12}\text{O}_{40}^{3-}$ precipitation occurs in the presence of C_8E_4 micelles, preventing SAXS measurement. Nevertheless, precipitate formation hints at a particularly strong interaction of $\text{Na}_3\text{PW}_{12}\text{O}_{40}$ with the surfactant.

Overall, we showed herein that the Keggin-type POMs interact with the γ -CD through various recognition modes depending on the POM's chaotropic power. The different interaction types previously identified can be classified along the Keggin ion series related to the magnitude of the binding constants. As tentatively depicted in Fig. 7, lowering of the ionic charge gives rise to multimodal supramolecular processes consistent with our experimental observations. On the left side of the scale, the $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ ion shows nearly no ability to interact with γ -CD and the $[\text{BW}_{12}\text{O}_{40}]^{5-}$ species reveals only weak interactions that involve preferably the outer surface of the CD as observed in the crystal structure of $\text{BW}_{12}\cdot 2\text{CD}$ (see Fig. 2c). Lowering further the ionic charge down to $[\text{SiW}_{12}\text{O}_{40}]^{4-}$ ion leads to the formation of 1:1 and 1:2 inclusion complexes involving mostly the primary face featuring higher binding constants. Finally, on the right side, the lowest ionic charge anion $[\text{PW}_{12}\text{O}_{40}]^{3-}$ gives rise to the highest binding constants reflected in the large diversity of supramolecular binding modes. As shown previously by ^1H NMR, both primary and secondary faces contribute to the POM's complexation and complementary DOSY NMR study suggested even the presence of an unusual supramolecular assembly corresponding to a γ -CD sandwiched by two $[\text{PW}_{12}\text{O}_{40}]^{3-}$ anion. Despite the lack of any structural model for such an arrangement, NMR showed clearly that in excess of POM, the primary face represented by H5 and H6 resonances and the secondary face detected by H3 resonances are both simultaneously involved in low self-diffusion coefficient assembly compatible with a $\{\text{POM}\cdots\text{CD}\cdots\text{POM}\}$ arrangement. The binding on the primary and secondary faces of the CD, which results in a large contact surface with the POM, is favored over the binding of POMs on the exterior side of the CD.

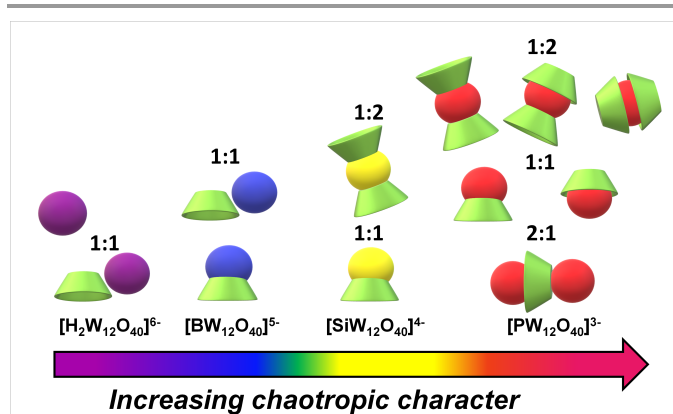


Fig. 7 Illustration of the most representative binding modes involved in the supramolecular assemblies of γ -CD (green torus) and the Keggin anions (spheres).

Conclusions

A comparative study allowed for an establishment of the relationship between the ionic charge of Keggin-type polyoxotungstate anions and their affinity towards a non-ionic substrate, γ -cyclodextrin. As the charge is lowered systematically keeping molecular size and shape, the ability to form supramolecular complexes increases continuously as was also observed for the interaction of the POMs with non-ionic surfactant micelles. The effects governing such supramolecular assemblies are suggested to be directly related to the solvation properties of the Keggin ions and to their propensity to shed hydration water upon formation of POM-substrate contact. While the most charged POMs, $[\text{H}_2\text{W}_{12}\text{O}_{40}]^{6-}$ and $[\text{BW}_{12}\text{O}_{40}]^{5-}$ showed weak to moderate affinity towards γ -CD, interacting mainly with its exterior wall, the lowest charge density species formed sandwich-type host-guest complexes involving either the primary face or the secondary face, featuring larger shared interfaces. Thus, in the Keggin-type series, the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion previously identified as a superchaotrope⁶⁰ exhibits a remarkable capacity to bind γ -CD through its two faces, thus presenting unique aggregation behavior. The host-guest adaptability allows maximizing contact interfaces through weak dispersion forces, highly amplified through a solvent effect arising from the pronounced superchaotropic nature of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion. This remarkable solution behavior of the $[\text{PW}_{12}\text{O}_{40}]^{3-}$ ion is not restricted to the systems involving cyclodextrins and the present study gives a better understanding for other self-assembly processes involving non-ionic substrates in water.^{21,59,61,62} Furthermore, POM/CD association has dramatic consequences on the POM's redox properties, i.e. a decrease in the standard redox potential. This effect can be understood as a partial removal of the hydration shell of the redox active ions arising from the presence of bulky γ -CD in the close vicinity of the POM. It is worth noting that opposite effects corresponding to a redox potential increase were commonly observed for interaction with cations such as protonation or ion-pairing with lithium cation.

Finally, the propensity of the Keggin-type anions to bind to γ -CD can be dramatically tuned by reversible electron transfer as demonstrated herein. This opens opportunities for the design of smart molecular materials like sensors, artificial muscles, and actuators, responsive to redox or possibly light stimuli.

Experimental section

Chemicals

All reagents were purchased from commercial sources and used without further purification. $\text{K}_{4.3}\text{Na}_{0.7}[\text{BW}_{12}\text{O}_{40}]\cdot 17\text{H}_2\text{O}$,⁶³ $\text{H}_5[\text{BW}_{12}\text{O}_{40}]\cdot 29\text{H}_2\text{O}$,⁶⁴ $\text{K}_4[\text{SiW}_{12}\text{O}_{40}]\cdot 6\text{H}_2\text{O}$,⁶⁵ $\text{Na}_{0.6}\text{H}_{3.4}[\text{SiW}_{12}\text{O}_{40}]\cdot 16\text{H}_2\text{O}$,⁶⁶ $\text{K}_3[\text{PW}_{12}\text{O}_{40}]\cdot 5\text{H}_2\text{O}$, $\text{Na}_{1.4}\text{H}_{1.6}[\text{PW}_{12}\text{O}_{40}]\cdot 12\text{H}_2\text{O}$, and $\text{H}_3[\text{PW}_{12}\text{O}_{40}]\cdot 14\text{H}_2\text{O}$ ^{65,67} were prepared according to literature procedures and checked by routine analyses prior using. The synthesis of $\text{Rb}_6[\text{H}_2\text{W}_{12}\text{O}_{40}]\cdot 11\text{H}_2\text{O}$ is described below. Solutions were prepared in MilliQ-water.

Materials characterization

Fourier Transform Infrared (FT-IR)

Spectra were recorded on a 6700 FT-IR Nicolet spectrophotometer, using diamond ATR technique. The spectra were recorded on non-diluted compounds and ATR correction was applied.

Energy-dispersive X-ray spectroscopy

EDS measurements were performed using a SEM-FEG (Scanning Electron Microscope enhanced by a Field Emission Gun) equipment (JSM 7001-F, Jeol). The measures were acquired with a SDD XMax 50 mm² detector and the Aztec (Oxford) system working at 15 kV and 10 mm working distance. The quantification is realized with the standard library provided by the constructor using L α lines.

Thermal gravimetric analysis (TGA)

To determine water and organic contents, a Mettler Toledo TGA/DSC 1, STAR[®] System apparatus was used under oxygen or nitrogen flow (50 mL min⁻¹) at a heating rate of 5 °C min⁻¹ up to 700 °C.

Single-Crystal X-ray diffraction analysis

Intensity data collections were carried out at T = 200(2) K with a Bruker D8 VENTURE diffractometer equipped with a PHOTON 100 CMOS bidimensional detector using a high brilliance μ S microfocus X-ray Mo K α monochromatized radiation (λ = 0.71073 Å). Crystals were glued in paratone oil to prevent any loss of crystallization water. Data reduction was accomplished using SAINT V7.53a. The substantial redundancy in data allowed a semi-empirical absorption correction (SADABS V2.10) to be applied, on the basis of multiple measurements of equivalent reflections. Using Olex2,⁶⁸ the structure was solved with the ShelXT⁶⁹ structure solution program using Intrinsic Phasing and refined with the ShelXL⁷⁰ refinement package using Least Squares minimization. Heavier atoms (W) for each structure were initially located by direct methods. The remaining nonhydrogen atoms were located from Fourier differences and were refined with anisotropic thermal parameters. Positions of the hydrogen atoms belonging to the γ -CDs were calculated and refined isotropically using the gliding mode. In the compound **PW₁₂@CD**, numerous free water molecules (H₂O) located inside the voids are disordered. Thereby the contribution of solvent-electron density was removed using the SQUEEZE routine in PLATON, producing a set of solvent-free diffraction intensities. Crystallographic data for single-crystal X-ray diffraction studies are summarized in Table S1.

Isothermal Titration Calorimetry (ITC)

Formation constants and inclusion enthalpies were simultaneously determined for each γ -CD/POM system by the use of an isothermal calorimeter (ITC200, MicroCal Inc., USA). Degassed deionized water solutions were used in both cell (V_0 = 202.8 μ L) and syringe (40 μ L). After addition of an initial

aliquot of 2 μ L, 10 aliquots of 3.5 μ L of the syringe solution were delivered over 7 s for each injection. The time interval between two consecutive injections was 70 s, which proved to be sufficient for a systematic and complete return to baseline. The agitation speed was set to 1000 rpm. The resulting heat flow was recorded as a function of time. ITC titrations were realized at three temperatures (288, 298 and 308 K), with 5 mM γ -CD solution in the syringe and 0.25 mM POM solution in the cell. Values and uncertainties of formation constants and inclusion enthalpies were determined by global analysis of the binding isotherms, by means of a dedicated homemade program,⁷¹ implementing 1:1 and 1:1+1:2 binding polynomials.

Electrochemistry

Purified water was used throughout. It was obtained by passing water through a RiOs 8 unit followed by a Millipore-Q Academic purification set. All reagents were of high-purity grade and were used as purchased without further purification. Cyclic voltammetric (CV) experiments were carried out with an Metrohm Autolab PGSTAT12 potentiostat/galvanostat associated with a GPES electrochemical analysis system (EcoChemie). Measurements were performed at room temperature in a conventional single compartment cell. A glassy carbon (GC) electrode with a diameter of 3 mm was used as the working electrode. The auxiliary electrode was a Pt plate placed within a fritted-glass isolation chamber and potentials are quoted against a saturated calomel electrode (SCE). The solutions were deaerated thoroughly for at least 30 minutes with pure argon and kept under a positive pressure of this gas during the experiments.

Nuclear Magnetic Resonance (NMR)

All solution NMR spectra were measured in D₂O at 26 °C. ¹H NMR spectra were obtained on a Bruker Avance 400 spectrometer at a Larmor frequency of 400.1 MHz, using 5 mm standard NMR tubes. The 1D ¹H spectra were recorded with one pulse sequence at 30° flip angle (pulse duration 2.7 μ s), using 0.1 s recycle delay, 3 s acquisition time, and 8 number of scans. Translational diffusion measurements were performed using Bruker's "ledbpgs2s" stimulated echo DOSY pulse sequence including bipolar and spoil gradients. Apparent diffusion coefficients were obtained using an adapted algorithm based on the inverse Laplace transform stabilized by maximum entropy.⁷² Chemical shifts are reported relative to tetramethylsilane (TMS).

Cloud point measurements

Aqueous solutions containing 60 mM C₈E₄ and a given amount of POM were prepared into 1 mL glass vials. The vials were placed in a temperature controlled water bath. The heating rate was constant at 1 K.min⁻¹ and the CP was detected by visual inspection after occurrence of turbidity. The choice of the counter-ion has negligible effects on the cloud point evolution.²²

Syntheses

Rb_{4.5}Na_{1.5}[H₂W₁₂O₄₀]•11H₂O

The synthetic procedure is inspired from the syntheses of the tetraalkylammonium (TAA) salts, TAA₆[H₂W₁₂O₄₀]•9H₂O.⁷³ Na₂WO₄•2H₂O (10 g, 30 mmol) is dissolved in 30 mL of deionized water, and the resulted solution is acidified with HCl (6 mol.L⁻¹) till pH ~ 3. The solution is kept boiling for 24 h and the pH remains at 3. After cooling to room temperature, the solution is then filtered, and rubidium chloride (5.5 g, 45 mmol) is added. The white solid product is collected, and finally dried in air. Yield: ~ 3.8 g (33%). IR (cm⁻¹): 956 (sh), 935 (s), 896 (s), 878 (s), 766 (vs), ¹⁸³W NMR: -117.6 ppm; ¹H NMR: 5.99 ppm. ICP-OES Anal. Calcd for H₂₄Na_{1.5}O₅₁Rb_{4.5}W₁₂: Na, 1.00; Rb, 11.10; W, 63.66. Found: Na, 0.9; Rb, 10.5; W, 58.0. TGA showed a weight loss of 5.5% in the 20-220 °C temperature range corresponding to the eleven hydration water molecules (calculated 5.7%).

CsK₂H₂[(BW₁₂O₄₀)•2(C₄₈H₈₀O₄₀)]•29H₂O (BW₁₂•2CD)

A mixture of K_{4.3}Na_{0.7}[BW₁₂O₄₀]•17H₂O (68 mg; 0.02 mmol) and γ-CD (85 mg; 0.06 mmol) in 3.5 mL aqueous solution of CsCl (0.05 mol.L⁻¹) and KCl (0.1 mol.L⁻¹) was stirred until obtaining a clear solution. The solution was then allowed to stand for crystallization in air. Colorless crystals of BW₁₂•2CD were obtained by slow ethanol vapor diffusion into water solution (10 mL vessel containing 2 mL of water solution was placed into a 250 mL vessel with a tight-fitting lid containing 50 mL of ethanol) after 3 days, and were collected, washed with mixture water/ethanol and dried on air. Yield 37 mg, 30%. FT-IR spectrum is given in Fig. S23. ICP-OES Anal. Calcd for BC_{1.5}CsH₂₂₀K₂O₁₄₉W₁₂: B, 0.17; C, 18.64; K, 1.26; W, 35.66. Found: B, 0.2; C, 18.5; K, 1.2; W, 32.5. EDS showed Cs : K : W ratio = 1.7 : 2.1 : 12. TGA showed a weight loss of 8.4% in the 20-220 °C temperature range corresponding to the twenty-nine hydration water molecules (calculated 8.4%) and a weight loss of 42% in the 220-700 °C range assigned to the loss of 2 CD (calculated 41.1%).

Na₃[(PW₁₂O₄₀)•(C₄₈H₈₀O₄₀)]•10H₂O (PW₁₂@CD)

Na_{1.4}H_{1.6}[PW₁₂O₄₀]•12H₂O (0.3 g; 0.096 mmol) is dissolved in 10 mL aqueous solution of NaCl (0.2 mol.L⁻¹) and γ-CD (0.138 g; 0.096 mmol) was added. The clear solution is stirred for 15 min. The solution was then allowed to stand for crystallization in air. Needle-like colorless crystals of PW₁₂@CD suitable for single crystal X-ray diffraction appeared within 5 days. They were isolated by filtration and washed with cold water. Yield 0.35 g, 76%. FT-IR spectrum is given in Fig. S24. ICP-OES Anal. Calcd for C₄₈H₁₀₀Na₃O₉₀PW₁₂: C, 13.03; Na, 1.56; P, 0.70; W, 49.87. Found: C, 12.8; Na, 1.6; P, 0.7; W, 48.9. EDS showed Na : P : W ratio = 2.7 : 1.1 : 12. TGA showed a weight loss of 3.9% in the 20-220 °C temperature range corresponding to the ten hydration water molecules (calculated 4.1%) and a weight loss of 26.8% in the 220-700 °C range assigned to the loss of 1.3 CD (calculated 27%).

Conflicts of interest

There are no conflicts of interests to declare.

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