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Synthesis of $\text{GdBO}_3:\text{Eu}^{3+}$ nanophosphors via acetylacetonate assisted aqueous sol–gel route: Effect of some synthesis parameters

 The corrections made in this section will be reviewed and approved by a journal production editor.

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Abstract

Nano-sized Eu^{3+} -doped GdBO_3 vaterite type nanophosphors have been synthesized successfully via acetylacetonate assisted aqueous sol–gel route. Systematic investigation of the influence of Eu^{3+} ion content and pH value of the suspension of precursors on structural, morphological and luminescence spectroscopic properties has been performed. X-ray powder diffraction (XRD), Fourier transform infrared spectroscopy (FTIR), Raman spectroscopy, field emission-scanning electron microscopy (FE-SEM), Thermogravimetric and differential thermal gravimetry (TG/DTA), room temperature steady and time resolved photoluminescence (PL) spectroscopy techniques were employed to characterize the obtained powder samples. It has been found that the Eu^{3+} ion content and pH value precursors suspension have important effect on the structural, morphology and emission properties. The XRD analysis reveals that $\text{GdBO}_3: 10\% \text{ mol Eu}^{3+}$ synthesized at $\text{pH} = 2$ and $\text{pH} = 10$ exhibit the quasi-purest $P6_3/mmc$ vaterite structure. The vibrational absorption of $(\text{BO}_3)^{3-}$ group is assigned to the GdBO_3 with calcite type or other vaterite structure namely triclinic and monoclinic. The morphologies of GdBO_3 particles size increases with increasing of pH value. All sample present only the orange-red colors emission from $^5\text{D}_0$ levels to the $^7\text{F}_J$ levels ($J = 0-3$) of Eu^{3+} ion. Furthermore, it was found that 10% concentration of Eu^{3+} presents the quenching of emission in GdBO_3 . Also, it was that the emission intensity was affected by the pH value and sample with $\text{pH} = 10$ exhibits the highest emission intensity. The asymmetry ratio parameter R between red

and orange emission intensities in function of Eu^{3+} content and pH value is determinate and discussed.

Keywords: $\text{GdBO}_3:\text{Eu}^{3+}$; Phosphors; Aqueous sol–gel; Photoluminescence

1 Introduction

Synthesis of rare earth ions doped inorganic nanophosphors with controllable, very clear morphologies has been an important challenge of modern materials physics and chemistry. Actually, it is well established that the properties of such luminescent nano-crystals depend on their structure, phase, dimensionality, shape, and size distribution in addition to their chemical composition. Actually, the industry demand is still very important in this field of advanced innovative materials due to their potential applications in many luminescent devices such as scintillators [1–3], Plasma Display Panels (PDP) [4–13] and Hg-free fluorescent lamp [14], [Light-emitting diode LEDs](#) [15]. Among these materials, REBO_3 Orthoborates have attracted much more attention. Thanks to their special optical properties, low synthesis temperature, low toxicity, high stability, large band gap, strong vacuum ultraviolet (VUV) absorption and high ultraviolet and optical damage threshold [16–19]. Up until now, there is no consensus about the crystallographic structure of vaterite and several structural with different space groups have been published such as: a disordered ($P6_3/mmc$) and an ordered hexagonal ($P6_3/mcm$) structure [20], $P6_3/c2$ space group [21], monoclinic (C_2/c) pseudo-vaterite-type structure [22], pseudo-wollastonite-type structure to Morganet.al. [23], hexagonal ($P6_3/m$) [24] and recently (C_2/c) monoclinic and P_1 triclinic phases [25].

Some red orthoborates phosphors such as $\text{YBO}_3:\text{Eu}^{3+}$ is actually used widely as commercial phosphors. Because of its poor chromaticity, due to the dominant orange $^5D_0 \rightarrow ^7F_1$ transition compared to red $^5D_0 \rightarrow ^7F_2$ transition, $\text{REBO}_3:\text{Eu}^{3+}$ does not make the requested VUV red luminescent material [26]. In fact, the red $^5D_0 \rightarrow ^7F_2$ transition is very sensitive to the symmetry of the crystal field around Eu^{3+} ions and its emission intensity is high if the symmetry is low. It is already suggested that the problem of the low chromaticity, can be resolved by lowering of symmetry of the crystal field. Therefore, the position of Eu^{3+} ion in $\text{REBO}_3:\text{Eu}^{3+}$ nanocrystals Orthoborates strongly affects their spectroscopic properties. Several methods have been performed to reduce the Eu^{3+} site symmetry: (i) by developing of such material with different crystal structure, in which Eu^{3+} ions occupy sites with low-symmetry environment [25], (ii) by synthesizing of nanoscaled $\text{REBO}_3:\text{Eu}^{3+}$ phosphors. In fact, by decreasing of the crystallite size, the number of Eu^{3+} ions occupying the surface with distortion of inversion symmetry can be increased. Therefore, using nanocrystals phosphors could be improved chromaticity [27–29]. Furthermore, it is well known that the method of synthesis has an important influence on the properties of prepared nanomaterials [30]. Indeed, $\text{GdBO}_3:\text{Eu}^{3+}$ micro/nanophosphors have been synthesized by several various methods, such as solid-state reaction, hydrolysis, ultrasonic spray cost alternative approach to synthesizing novel nanoscale inorganic materials activated by the rare earth ions in powders and films forms. This method have becomes very popular due to its versatility, better purity, low processing temperature, and the possibility of controlling the size and morphology of particles. In addition, the doped nanomaterials prepared by sol–gel method are characterized by their ability to provide uniform distribution of the dopant within the host phase. Moreover, within the sol–gel method, different conditions of synthesis influencing the final products are very important parameters such as complexing and chelating agents, dopant concentration, pH value of solution and annealing temperature. The pH value of the solution plays a crucial role and its change can cause some alteration in structural, the morphological and spectroscopic properties [31,32].

In this work, GdBO₃:Eu³⁺ orthoborates vaterite type nanophosphors have been synthesized by a simple aqueous sol-gel method using the acetylacetonate (CH₃COCH₂COCH₃) as chelating agent. A systematic study of the effect of only two synthesis parameters namely the Eu³⁺ ion content (0.5, 1, 2, 3, 10, 15 and 20 mol %) and pH value precursors suspension, extending from acid to basic states (pH = 2, 4, 7, 8 and 10) on the structural, morphology and photoluminescence spectroscopic has been performed. In this study, we begin to fix the pH value (pH = 2) and we optimize the value of Eu³⁺ content and after that the pH value is optimized. Thus, structural purity and luminescence properties are the criteria for selecting the best parameter. Furthermore, in this work, we demonstrate that both Eu³⁺ ion content and the pH value play a significant role in the determination of the phase purity and the photoluminescence properties of GdBO₃:Eu³⁺ nanoparticles.

2 Experimental details

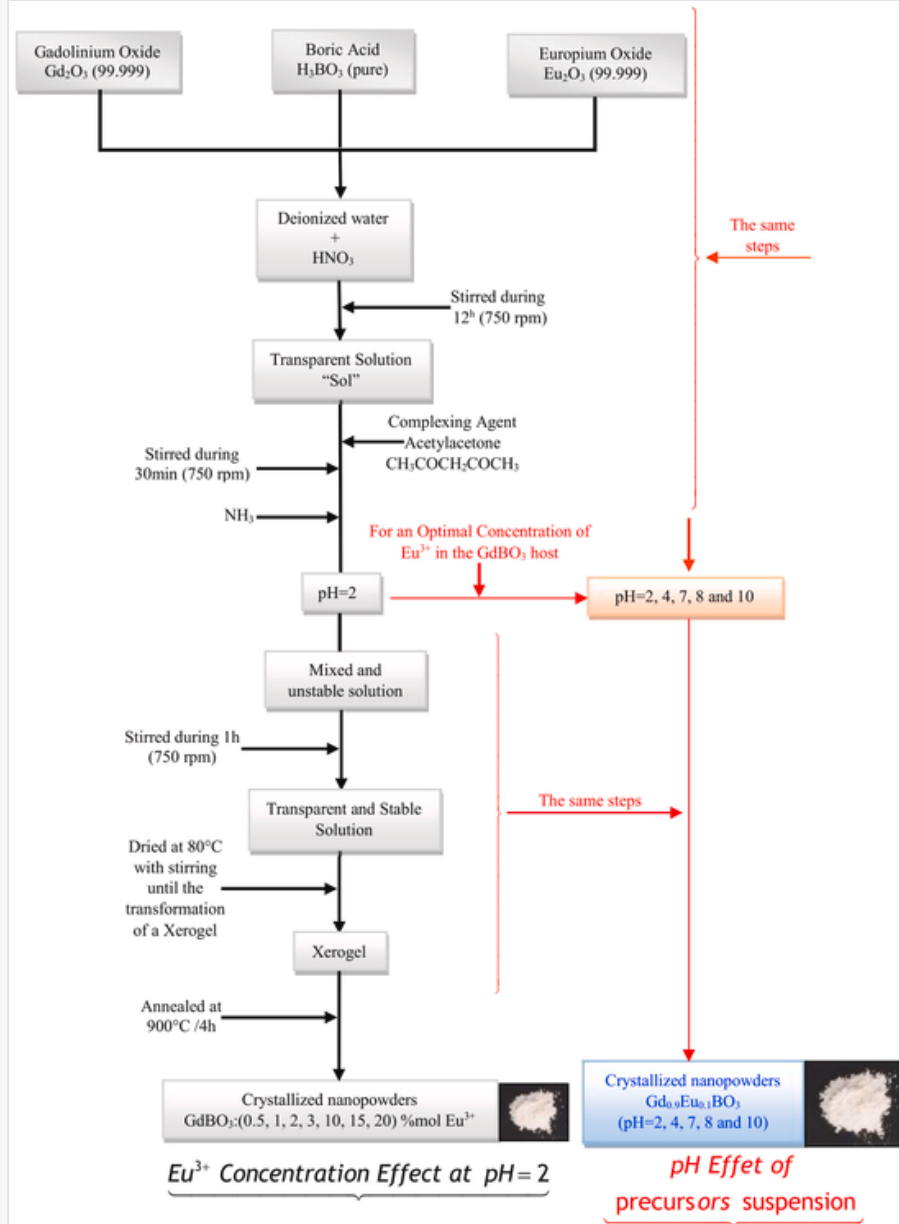
2.1 Samples preparation

GdBO₃:Eu³⁺ nanopowders were prepared by aqueous sol-gel method. First, in order to determine the quenching concentration of Eu³⁺ ions in GdBO₃ host, the GdBO₃ was doped with various concentration of Eu³⁺ with its general formula Gd_{1-x}Eu_xBO₃ (x = 0.5, 1, 2, 3, 10, 15 and 20) at pH = 2. For getting this material, stoichiometric amounts of high purity gadolinium oxide Gd₂O₃ (99.999%), europium oxide Eu₂O₃ (99.999%) and boric acid H₃BO₃ were dissolved in de-ionized water by the addition of nitric acid (HNO₃). In order to compensate the volatilization of B₂O₃ at high temperature during the heating treatment, an excess of 15% mol of H₃BO₃ was added into the raw materials. The mixed solution was stirred at room temperature for 12h (750 rpm). Then, a transparent solution was obtained and an amount of acetylacetonate (CH₃COCH₂COCH₃) corresponding to 1 mL for 1g of product was added as complexing agent (Fig. 1). The chemical reaction can be described with stoichiometric equation following:



alt-text: Fig. 1

Fig. 1



Schematic diagram of the synthesis Eu^{3+} doped $GdBO_3$ by aqueous sol-gel route.

To stabilize the solution, stirring was carried out for 30 min. After that, the pH value of precursors suspension was adjusted by adding ammonium hydroxide NH_4OH at pH = 2. This solution was then stirred an extra hour and a homogeneous and transparent aqueous solution was obtained. The resulting solution was dried at 80 °C

until transforming in to a Xerogel. Finally, the Xerogel were introduced in a furnace and heated in air at 900 °C for 4 h for different Eu^{3+} concentration. Secondly, the preparation of GdBO_3 doped with the Eu^{3+} obtained optimal concentration at different pH of precursors suspension namely pH = 2, 4, 7, 8 and 10 has been done by the same steps as we saw in the first part of this synthesis (Fig. 1).

2.2 Characterization

The phase identification and the related properties of the nanopowders were investigated by X-ray diffraction (XRD) technique a PAN analytical X'Pert (Philips) PRO using CuK_α radiation ($\lambda = 1.54059 \text{ \AA}$) operated at 45 kV and 40 mA. A symmetric (θ - θ) scans were performed from 10° to 90° 2θ with a step width of 0.02° . All the data were processed by X'Pert High Score plus Software with commercial databases (FWHM deduction and peak identification). Rietveld refinement was done with help of the FullProf software. Thermogravimetric analysis (TG) and differential thermal gravimetry (DTA) were measured by SETARAM SetsysEv 1750 (TGA-DTA 1600) instrumentation. Micrographs were investigated by means of a ZEISS Supra 55VP Scanning Electron Microscope (SEM) operating in high vacuum at 3 kV, using secondary electron detector (Everhart-Thornley detector). Specimens were prepared by sticking powder onto the surface of an adhesive carbon film. Raman spectra were recorded at room temperature under excitation at 514.5 nm supplied by a Spectra Physics-Stabilite 2017-100 mW argon laser, using a T64000 Jobin-Yvon spectrometer equipped with a CCD detector cooled with liquid nitrogen. For all the samples, the acquisition time was fixed at 30 s and three measurement cycles were performed. The infrared spectra were recorded in the range of $400\text{--}4000 \text{ cm}^{-1}$ with a Nicolet-IR 380 Fourier transform infrared (FTIR) spectrometer. The steady and time-resolved photoluminescence spectra were carried out using PerkinElmer (LS-50B) luminescence spectrometer with pulsed Xe lamp excitation at room temperature as described in [33].

3 Experimental results and discussion

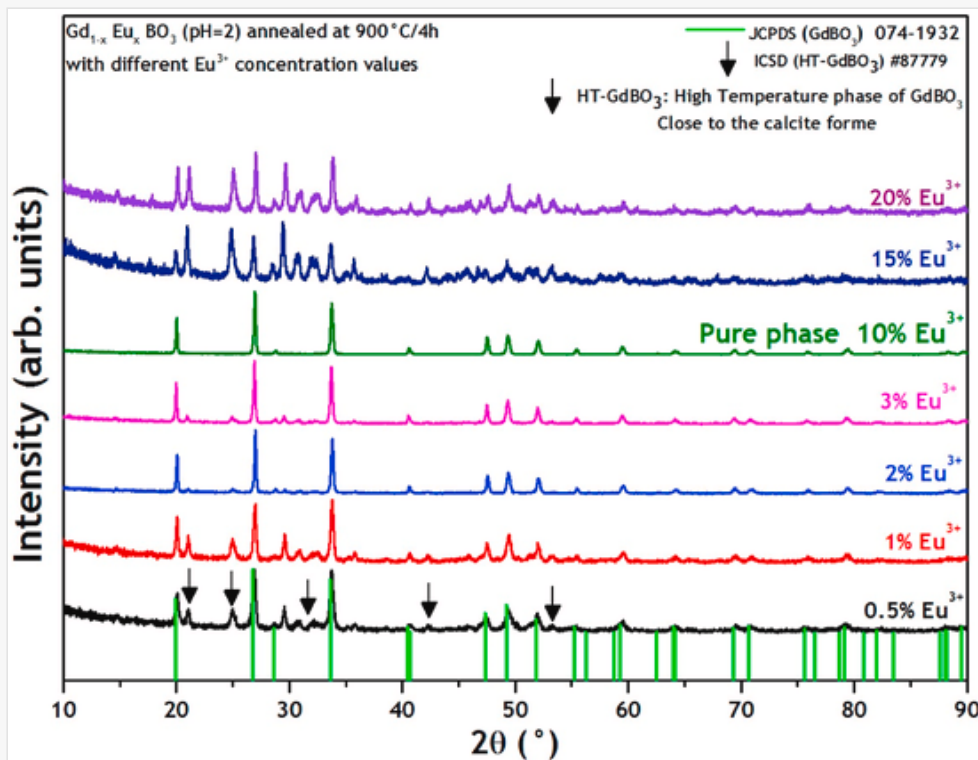
3.1 Structural analysis

Fig. 2(a) shows the powder XRD patterns of the synthesized $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ ($x = 0.005, 0.01, 0.02, 0.03, 0.1, 0.15$ and 0.2) at pH = 2 with 15% excess of H_3BO_3 , annealed at 900 °C for 4h duration. All diffraction peaks in Fig. 2 was indexed according to the hexagonal phase of GdBO_3 as listed in JCPDS files (No. 74-1932). In this system, it is found that crystallographic structure purity is highly dependent on the Eu^{3+} concentration and introducing the Eu^{3+} dopant drastically changes the host matrix structure. In fact, from the patterns of Fig. 2(a), one can easily observe that it exist three concentration regions, according to the existence of the parasite phases, especially, High temperature phase (HT- GdBO_3), indicate by down arrow. The evolution of parasitic phases with Eu^{3+} content is as follows: a first is the low Eu^{3+} concentration region, from 0.5 at % to that of 3 at %, in which the parasite phase peaks are present. The second region around 10 at %, in which the parasite phase peaks vanished completely and GdBO_3 give a quasi-purest vaterite phase. From 15 to 20 at %, constitute the third region, in which the parasitic phases reappear, which become more intense. Although, all samples are prepared in the same conditions, it did not prevent this highly dependence of GdBO_3 phase on Eu^{3+} content. Usually the parasitic phases are not observable for low concentrations of the dopant, but can appear for high concentrations. In the present work is the opposite was observed, indicating that there are other factors involved and their effects is more important than that of the concentration. This phenomenon has been observed in orthoborates LnBO_3 prepared by solid state reaction [26]. Some authors have already noticed the dependence of the phases produced as a function of the concentration of the dopant. Szczeszak et al. [34] found that in $\text{La}_{1-x}\text{Eu}_x\text{BO}_3$ and

$\text{La}_{1-x}\text{Tb}_x\text{BO}_3$ nanopowders synthesized by sol-gel Pechini method the aragonite phase purity depend on the Eu^{3+} and Tb^{3+} concentration, which connected just to the annealing temperature.

alt-text: Fig. 2(a)

Fig. 2(a)



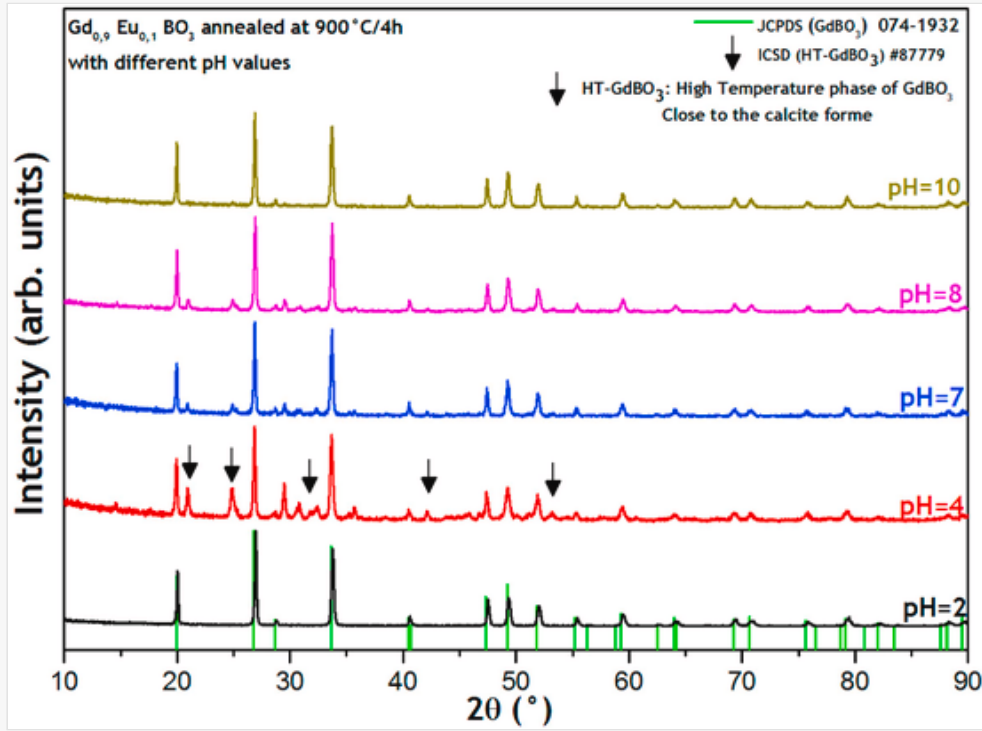
. XRD patterns of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ ($\text{pH} = 2$) annealed at $900\text{ }^\circ\text{C}$ for 4 h synthesized by aqueous sol-gel route.

The sample with 10 at % Eu^{3+} content presents an apparent pure hexagonal vaterite phase compared to the other concentration and it was chosen for the study of the pH value effect.

Fig. 2(b) shows the powder XRD patterns of the $\text{GdBO}_3: 10\text{ at } \% \text{Eu}^{3+}$ prepared with 15% excess of H_3BO_3 and prepared at different pH, ranging from $\text{pH} = 2$ to 10, annealed at $900\text{ }^\circ\text{C}$ for 4h. As it was observed that for both $\text{pH} = 2$ and $\text{pH} = 10$, the samples exhibit a pure hexagonal GdBO_3 vaterite structure. For the other pH values (4, 7 and 8), parasitic phases are also observed, such as HT- GdBO_3 , indicated by down arrow. This demonstrated that the pH value has an effect on the purity of GdBO_3 crystallographic structure. Indeed, the explanation of the mechanisms responsible for the formation of the structural phases produced is very complicated. In fact, the B_2O_3 is much more volatile than the gadolinium and europium oxides leading to stoichiometric imbalance [26]. Furthermore, it is not possible to control the amount of B_2O_3 which can volatilize during the synthesis in a random manner. This stoichiometric deviation can influences the reaction rate of gadolinium oxide with europium and boron oxides, which can produces additional phase traces such as gadolinium oxides (Gd_2O_3), borates oxide (Gd_3BO_6) [26] and/or other possible parasitic unknown phases. Consequently, from the above, we can say that the formation of $\text{GdBO}_3: x\text{ at } \% \text{Eu}^{3+}$ in it pure vaterite phase and/or others additional phases depends on all oxides concentration present in real time (B_2O_3 , Gd_2O_3 and Eu_2O_3) during the reaction and depends also to the used synthesis method (pH value, annealing temperature, duration, reagent..).

alt-text: Fig. 2(b)

Fig. 2(b)



XRD patterns of $Gd_{0.9}Eu_{0.1}BO_3$ with different pH value of precursor suspension, annealed at 900 °C for 4 h synthesized by aqueous sol-gel route.

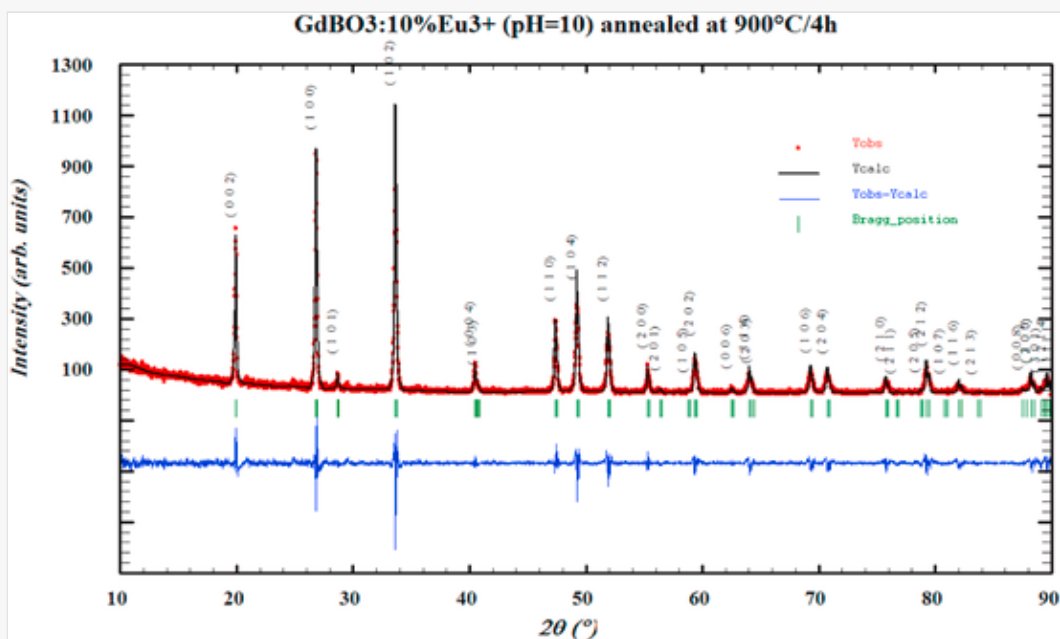
The Rietveld analysis was used to accurately determine the structure and the crystallographic parameters by using Fullproof suite program for all samples as indicated in Fig. 2(c). On this figure we indicate the calculated and observed diffraction profile of $GdBO_3$ doped at 10 at% Eu^{3+} at pH = 10 delivered by the Rietveld refinements as example in hexagonal $P6_3/mmc$ space group. The line $Y_{obs} - Y_{calc}$ represents the difference between the calculated and observed intensity. The crystallographic data and atomic parameters for $GdBO_3$ doped at 10 at% Eu^{3+} at pH = 10 are summarized in Table 1. The crystal lattice strain generated by the annealing temperature is determined from the Williamson–Hall (W–H) relationship [35]:

$$\beta \times \frac{\cos \theta}{\lambda} = \frac{1}{D} + \eta \sin \theta / \lambda$$

(1)

Where β is the full width at half maximum, λ is the X-ray wavelength, θ is the diffraction angle, D is the effective crystallite size and η is the effective strain. In Fig. 3(a) and Fig. 3(b), we present the variation of the crystallite size as well as the strain quantities against the Eu^{3+} ion content (Fig. 3(a)) and pH value (Fig. 3(b)) in $GdBO_3$ host material. At 10 at % Eu^{3+} , the sample presents the biggest crystallite size. From Fig. 3(a), it is observed that the below 10% mol Eu^{3+} , both crystallite size and strain increases from 0.5 to 10% mol, after that they evolve in the opposite direction. The strain follows the same behavior than that of grain size. Both crystallite size and strain behaviors can relate to the phase parasites in addition to the defects, which can exist in the samples. Furthermore, with the variation of pH value of solution the $GdBO_3$: 10%mol Eu^{3+} crystallite size change (Fig. 3(b)). The crystallite size increases from pH = 2 to pH = 4 and decreases after from pH = 4 to pH = 10. Although that the sample with pH = 4 presents some parasite phase (Fig. 3(b)) it exhibits the biggest crystallite size and with the highest strain value.

Fig. 2(c)



The Rietveld profile fits of X-Ray diffraction data of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ (pH = 10), annealed at 900 °C for 4 h.

alt-text: Table 1

Table 1

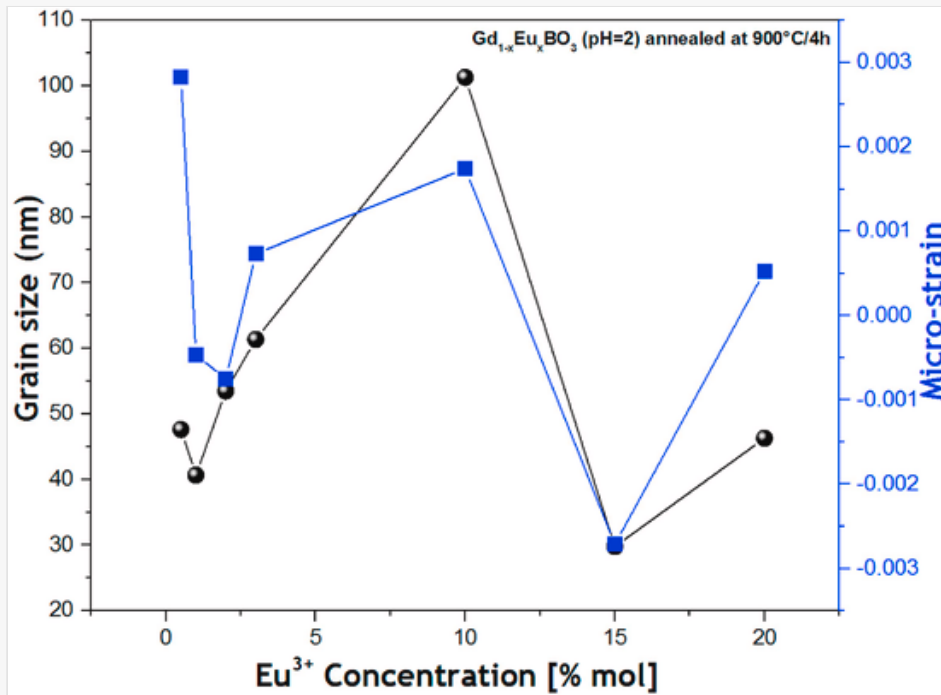
i The table layout displayed in this section is not how it will appear in the final version. The representation below is solely purposed for providing corrections to the table. To preview the actual presentation of the table, please view the Proof.

Final refined structure parameters of $\text{GdBO}_3:\text{Eu}^{3+}$ samples using derived from the FullProf refinement of X-ray diffraction data.

Chemical formula	GdBO_3
Crystalsystem	Hexagonal
Space group	$P6_3/mmc$
a (Å) = b (Å)	3.8326 (6)
c (Å)	8.909 (2)
V (Å ³)	113.33
Crystallite size (nm)	74
R _{exp} %	26.91
R _{wp} %	36.6
R _p %	39.4
χ^2	1.85

alt-text: Fig. 3(a)

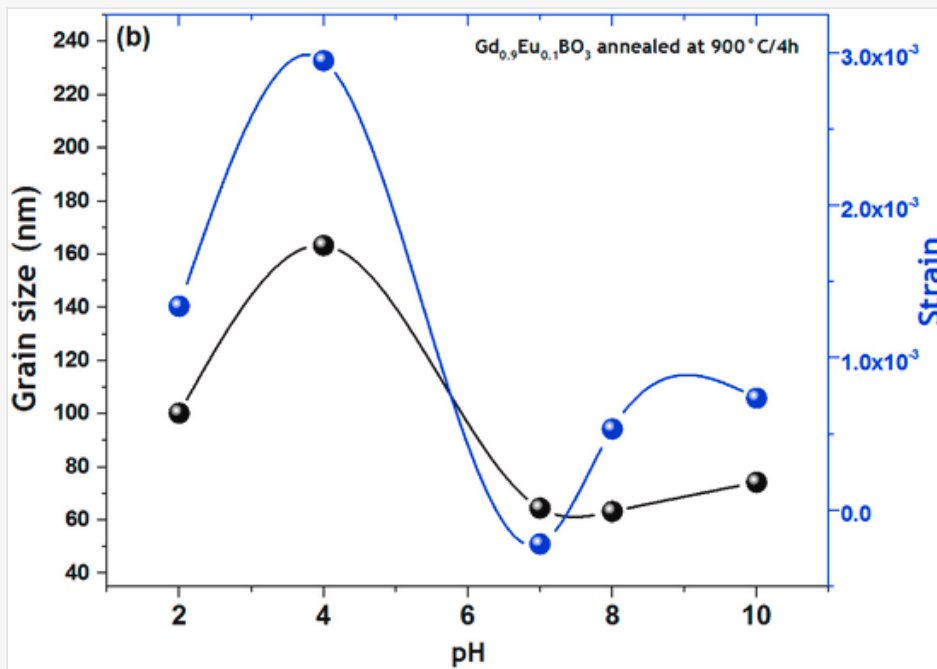
Fig. 3(a)



Crystallite size and microstrain evolutions of $Gd_{1-x}Eu_xBO_3$ ($pH=2$) powders, annealed at $900\text{ }^\circ\text{C}/4h$

alt-text: Fig. 3(b)

Fig. 3(b)



Crystallite size and microstrain evolutions of $Gd_{0.9}Eu_{0.1}BO_3$ powders with different pH value of precursors suspension, annealed at $900\text{ }^\circ\text{C}/4h$

Also, from Fig. 3(a) and (b), it is important to note, the existence of some experimental points coming out of the curves, located principally for the Eu³⁺ ion weak and high concentrations as well as the high pH values. One

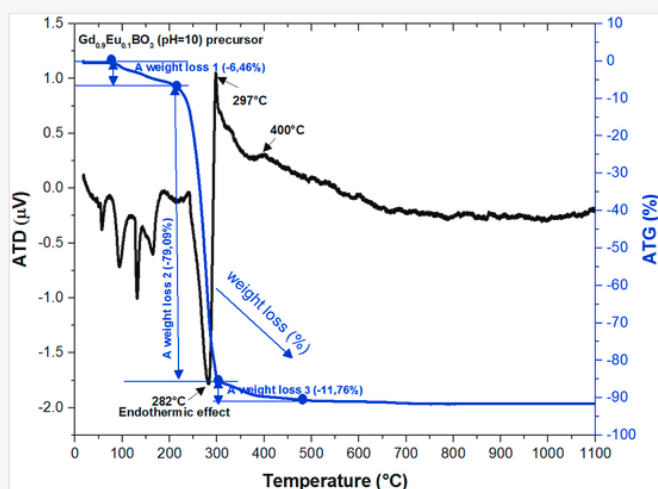
can think that these points can be related to the purity of the GdBO_3 vaterite phase, especially for the Eu^{3+} ion concentration parameter.

3.2 Thermal analysis

Thermogravimetric curve of $\text{GdBO}_3:10$ at % Eu^{3+} for $\text{pH} = 10$ precursor pointed in blue in Fig. 4 consists in three mass losses regions. The first mass loss was observed between 80°C and 210°C and is attributed to the elimination of absorbed species, such as alcohol molecules or water. The second important mass loss observed between 210°C and 300°C , corresponds to GdBO_3 condensation step, and is in parallel with a large endothermic peak (around 282°C) in the DTA curve. This endothermic effect is attributed to the removal of H_2O molecules. Otherwise, in the range of 210°C to 300°C , there is formation of inorganic compounds. A third observed mass loss is between 300°C and to 490°C depending on the formation of GdBO_3 and corresponds to the first GdBO_3 crystallization temperature range. This final mass loss is characteristic of ultimate formation of the crystallites growth of the GdBO_3 network and final disposal of residual OH groups. Above 600°C , no weight loss was observed. It may be noted that the total mass loss is large (almost -91%) occurred during the heat treatment to 600°C . From the recorded curves, we can deduce that the best crystallization temperatures, which were chosen, should be in excess of 1000°C . The formation of hexagonal structure GdBO_3 is observed above 600°C , which is also seen in the curve of the DTA as a nearly linear and stable curve between 600°C and 1100°C .

alt-text: Fig. 4

Fig. 4



Thermogravimetric analysis of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ ($\text{pH} = 10$) synthesized by aqueous sol-gel route.

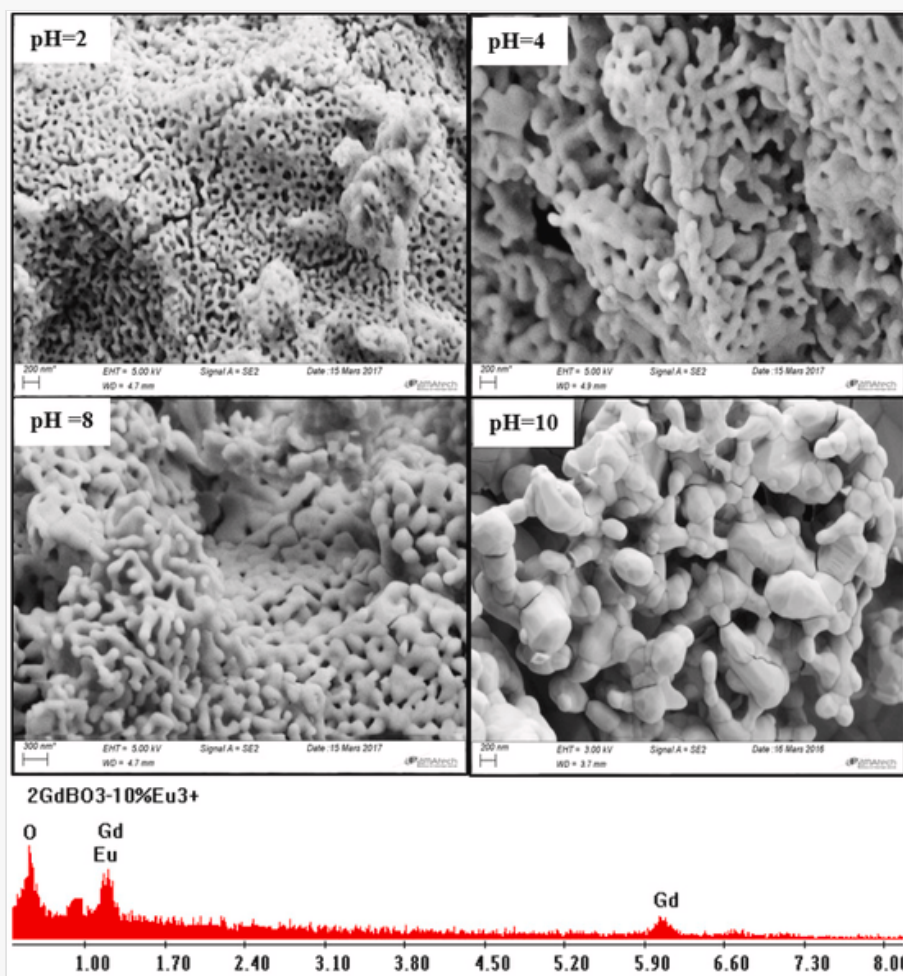
3.3 Morphology analysis

It is well known that the morphology and particle size of the phosphors are important for their in lighting devices applications. In Fig. 5 we show that the morphology and particle size of the $\text{GdBO}_3:10$ at % Eu^{3+} for $\text{pH} = 10$ powder for different pH value. From this figure, one can see that the morphology and size of the particles were dependent on the pH values. In fact, one can observe that the particles are slackly bound and connected to each other as networks with the particle size depend on the pH value. Indeed, for low $\text{pH} = 2$ the sample presents a small particle size, and slightly increases for $\text{pH} = 4$ and 8 . The sample prepared with $\text{pH} = 10$ exhibit a large particle size. Although, the crystallite size at $\text{pH} = 10$ is small, it has the largest particle size. The sizes of these primary particles were estimated about 150 and 250 nm, which can be considered as the agglomeration of three

to five crystallites estimated by W–H method [36]. EDAX analysis, clarifies the qualitative analysis of the elements of the respective phosphor composition i.e. Gd^{3+} , Eu^{3+} and O^{2-} . Furthermore, EDAX spectra elucidate the absence of any impurity elements, during sol gel reaction (Fig. 5).

alt-text: Fig. 5

Fig. 5



SEM morphology of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ (pH = 2, 4, 8 and 10) powders, annealed at 900 °C/4h synthesized by aqueous sol-gel route.

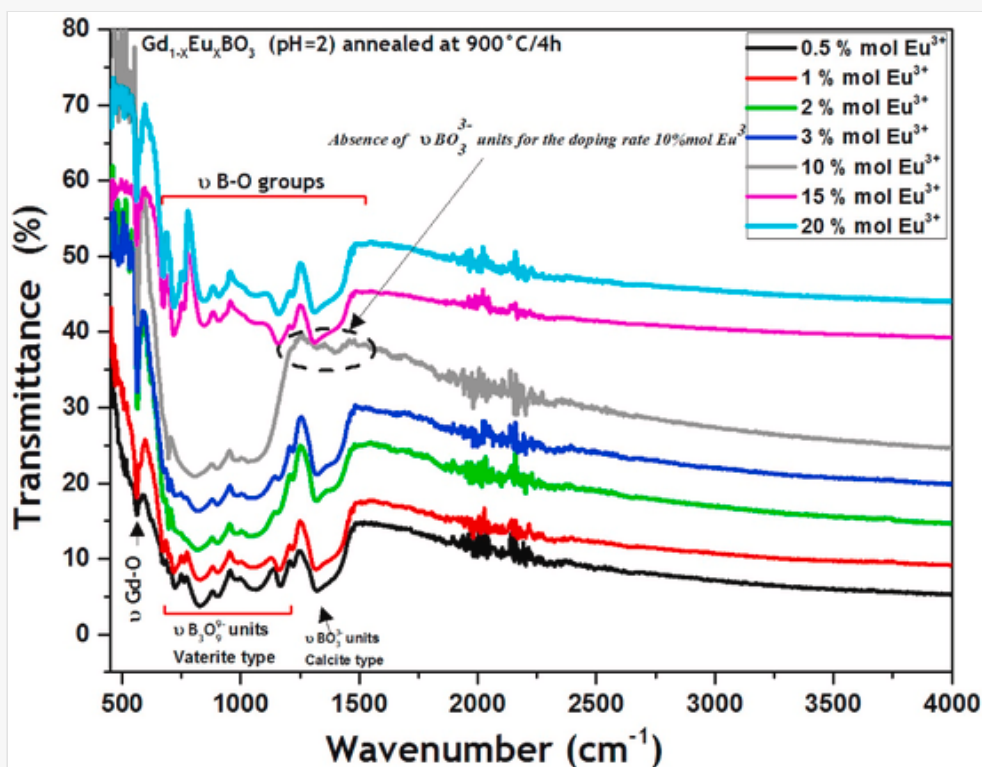
3.4 Vibrational analysis

It is well known that the trivalent rare-earth orthoborates presents three-isostructural forms, which are calcite, aragonite, or vaterite. Depending on the size of lanthanide cations, their orthoborates usually present the aragonite-type (La–Nd), vaterite-type (Sm–Yb), or calcite-type (Lu) structure. Furthermore, both two structures calcite-type and the aragonite-type take part triangular planar $(\text{BO}_3)^{3-}$ groups, but the vaterite-type borates participate $(\text{B}_3\text{O}_9)^{9-}$ groups. Thereby, the vibrational spectra of vaterite orthoborates type are found to be significantly different from those of calcite and aragonite types. Fig. 6(a) shows the FT-IR spectra of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ ($x = 0.005, 0.01, 0.02, 0.03, 0.1, 0.15, 0.2$, pH = 2 and 15% excess of H_3BO_3) annealed at 900 °C for 4h. As a first observation, the absorption bands and peaks of the water (H_2O) and OH molecules are absent. At low value of (pH = 2), the reaction is governed by (H_3O^+) ionic molecules in the solution and the reaction $\text{H}_2\text{O} + \text{H}^+ \rightarrow \text{H}_3\text{O}^+$ reduce the amount of the water in the final product [37]. Furthermore, one can observe that the FTIR vibrational spectra reflect the XRD analysis results. For all Eu^{3+} content, the samples exhibit the absorption bands in the region. $666\text{--}12250\text{ cm}^{-1}$, ascribed to the characteristic vibration modes of $(\text{B}_3\text{O}_9)^{9-}$

groups in the GdBO_3 vaterite orthoborates [12,38, and 39]]. The absorption intensity depends on Eu^{3+} concentration, which is related to the structure purity. In fact, for low Eu^{3+} content the samples present intense vibrational absorption, however, for high concentration (15 and 20 at % of Eu^{3+}) weak absorption. This difference can be related to the presence of phase impurity in GdBO_3 vaterite structure, generated by the high Eu^{3+} concentration. Furthermore, an intense absorption bands in $1225\text{--}1500\text{ cm}^{-1}$ region is observed for all sample except 10 at % Eu^{3+} . This band can be related to the asymmetric stretching of B–O in triangular $(\text{BO}_3)^{3-}$ groups in the calcite-type, with the analogy of FTIR spectra of calcite LuBO_3 orthoborates [38]. The FTIR spectra analysis shows that it exist some boron atoms are in 3-fold coordination and the $\text{GdBO}_3:\text{Eu}^{3+}$ structure made of the $(\text{BO}_3)^{3-}$ groups. To confirm the presence of GdBO_3 orthoborates in other structure together with vaterite type, two Raman spectra was registered (Fig. 6(b)) for the two Eu^{3+} content namely 3 and 10 at %. In fact, for 3 at % Eu^{3+} the sample holds the parasite phase and for 10 at % Eu^{3+} is pure. As shown in Fig. 6(b) for 10 at % Eu^{3+} content (red color) the Raman spectrum correspond to pure GdBO_3 vaterite structure [39]. The sample contains 3 at % Eu^{3+} presents some additional peaks mainly an intense one at 951.07 cm^{-1} and some weaks in $280\text{--}380\text{ cm}^{-1}$ region related to $(\text{BO}_3)^{3-}$ groups [39].

alt-text: Fig. 6(a)

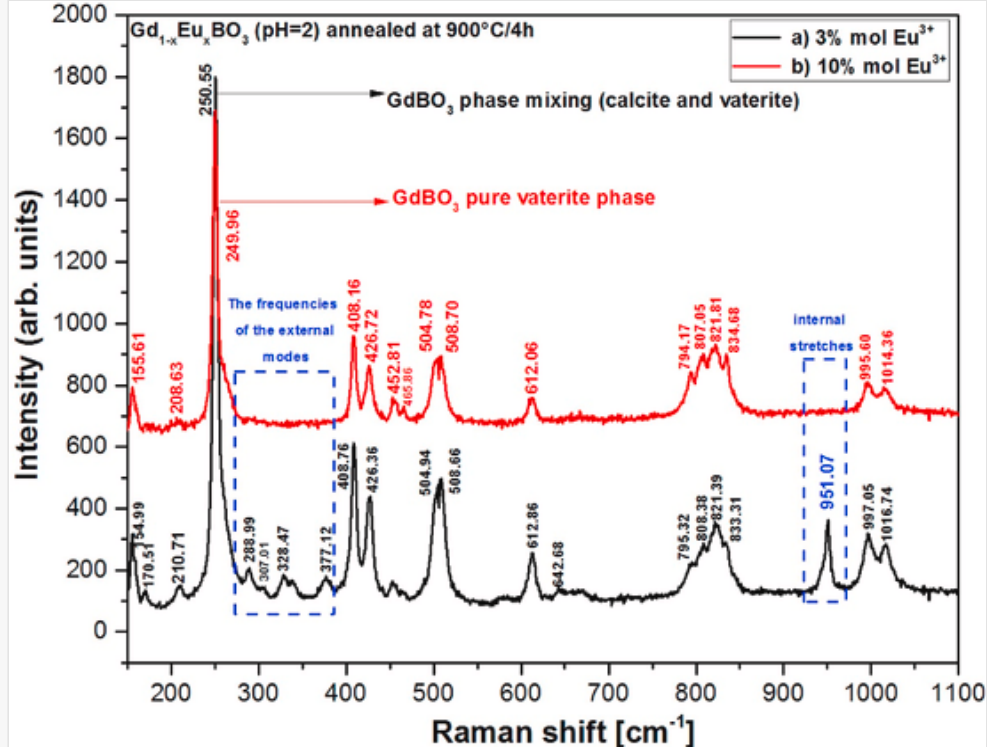
Fig. 6(a)



FTIR spectra of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ (pH = 2) annealed at 900°C for 4 h, synthesized by aqueous sol-gel route.

alt-text: Fig. 6(b)

Fig. 6(b)



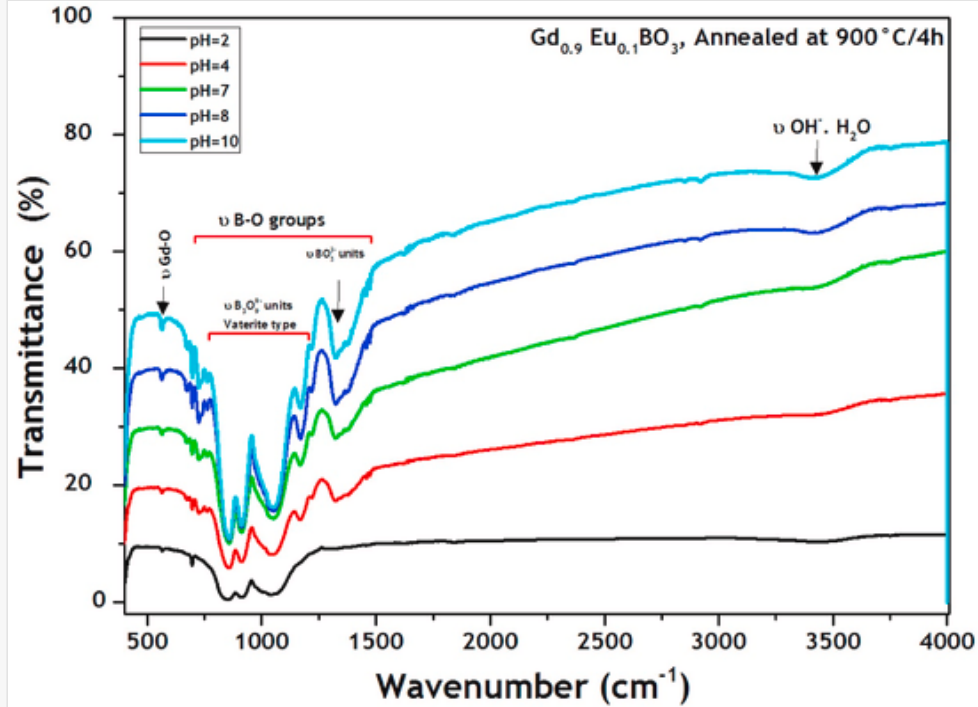
The Raman spectra of the $Gd_{1-x}Eu_xBO_3$: a) 3%mol Eu^{3+} , b) 10%mol Eu^{3+} , annealed at 900 °C for 4 h, synthesized by aqueous sol-gel route.

Earlier and in his work, Ren et al. [10] conclude that the high-temperature (HT) phase $GdBO_3$ crystallizes in a calcite, which consists of planar triangular borate group $(BO_3)^{3-}$. However, Szczeszak et al. [39] prepared new monoclinic $GdBO_3$ orthoborates by a modified Pechini method crystallized in the C_2/c space group. The same author Szczeszak et al. [25] has assigned the presence of borate group $(BO_3)^{3-}$ to new triclinic structure of $GdBO_3:Eu^{3+}$ with PT space group prepared by hydrothermal method. Furthermore, in our previous work [36], on the basis of the VUV excited emission, it is supposed that $GdBO_3:Eu^{3+}$ calcite type is presents as parasitic phase.

Fig. 6(c) shows the FTIR spectra for the as synthesized of $GdBO_3:10$ at % Eu^{3+} , annealed at 900 °C for 4 h samples prepared at different pH values. Same absorption bands characteristic of $(B_3O_9)^{9-}$ groups in $GdBO_3$ vaterite structure and $(BO_3)^{3-}$ group phase impurity discussed above. It is important to note that for pH = 2, absorption bands in 1225–1500 cm^{-1} region is not observed and which considered as a pure vaterite structure. One can note that the intensity of FTIR absorption bands at around 3500 cm^{-1} increases with increasing of pH values of solution. In fact, for high pH value the reaction is governed by (OH) molecules [37].

alt-text: Fig. 6(c)

Fig. 6(c)



FTIR spectra of $Gd_{0.9}Eu_{0.1}BO_3$ with different pH value of precursors suspension, annealed at 900 °C for 4 h, synthesized by aqueous sol-gel route.

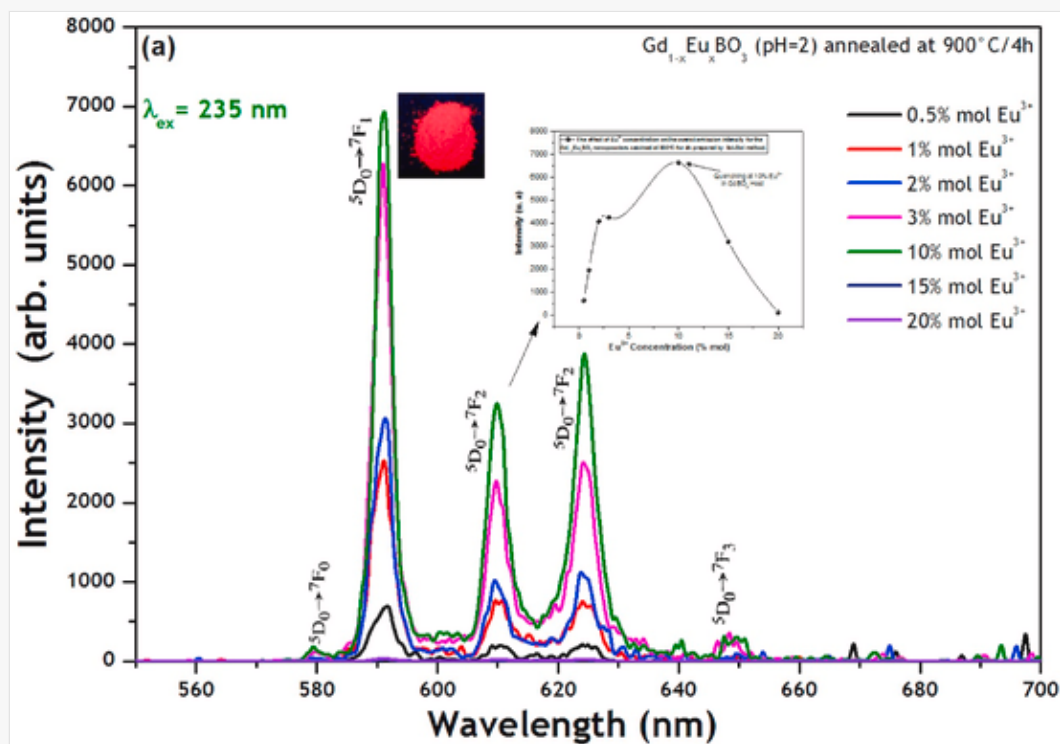
3.5 Photoluminescence spectroscopy

At the beginning, we have measured the room temperature emission spectra of $Gd_{1-x}Eu_xBO_3$ ($x = 0.005, 0.01, 0.02, 0.03, 0.1, 0.15, 0.2$, pH = 2 and 15% excess of H_3BO_3), annealed at 900 °C for 4 h duration, under 235 nm excitation wavelength (Charge transfer Band CTB) shown in Fig. 7(a). The spectra consist of sharp peaks ranging from 570 to 660 nm assigned the transitions from the excited 5D_0 level to the ground states of 7F_J ($J = 1, 2$ and 3) of the Eu^{3+} activator [40]. The major emission transitions are centered at 591 nm ($^5D_0 \rightarrow ^7F_1$), orange-red magnetic dipole (MD) and at 614, 624 nm ($^5D_0 \rightarrow ^7F_2$), which is the hypersensitive red electric dipole (ED). The weak peaks observed at 648 and 671 nm are ascribed to $^5D_0 \rightarrow ^7F_3$. The less intense emission peak at 578 nm is attributed to $^5D_0 \rightarrow ^7F_0$ forbidden transition. No intraconfigurational emission transitions from the highest excited states 5D_J ($J=1-4$) of Eu^{3+} or Gd^{3+} are observed. Furthermore, it is observed from Fig. 7 (a), the orange emission corresponding to the $^5D_0 \rightarrow ^7F_1$ magnetic dipole transition (591 nm) is greater than the red electric dipole transitions in the region ($^5D_0 \rightarrow ^7F_2$ situated at 614 nm and 624 nm) which is transitions from Eu^{3+} ions located in a non-inversion symmetry site. As shown in Fig. 7(a)inset, the concentration quenching occurs at 10 mol% of Eu^{3+} ions concentration, which is associated with the energy transfer, occurs from one Eu^{3+} ions to other ions and finally quenches the emission intensity [41]. Two effects can tuned the orange-red Eu^{3+} emission: (i) the existence of parasitic phases and the type of symmetry in which Eu^{3+} is located (centrometric or non-centrometric), (ii) the change in crystallite size. In fact, the reduction of size into nanoscale which increase the number of Eu^{3+} ions located in the surface area, leading to lowering the symmetry of crystal field and thereby increases the red emission ($^5D_0 \rightarrow ^7F_2$) relative to the orange one ($^5D_0 \rightarrow ^7F_1$). Therefore, the asymmetry ratio parameter R, which is the ratio of the integrated intensities of the luminescence transitions of $^5D_0 \rightarrow ^7F_2$ on $^5D_0 \rightarrow ^7F_1$ transitions, is introduced (Fig. 7(b)) [42,43]. In fact, this parameter can give

information about the Eu^{3+} ions local symmetry distortion in the host material. In addition, as a consequence of nanoscale materials, the value of R parameter is usually higher for the nanoscale material relative to its bulk form. From Fig. 7(b), one can observe that as the Eu^{3+} concentration increases from 0.5 to 3 at % Eu^{3+} , the R parameter increases and roughly stabilized between 3 and 10 at % Eu^{3+} to come back increases after that. The increasing in R parameter is related to the Eu^{3+} site symmetry lowering which can be generated by existence of parasite phases or defects. For the constant quasi-constant value of R is connected to the GdBO_3 hexagonal phase purity. Although, this low symmetry can gives better chromaticity but with lower emission intensity.

alt-text: Fig. 7(a)

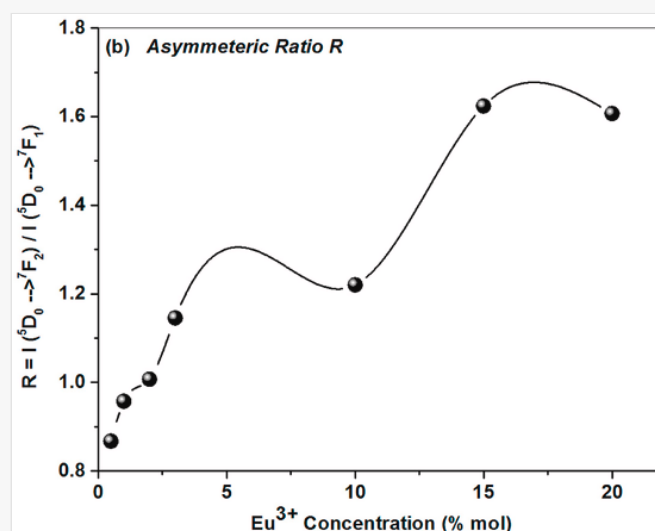
Fig. 7(a)



Emission spectra of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ (pH=2) powders, annealed at $900^\circ\text{C}/4\text{h}$

alt-text: Fig. 7(b)

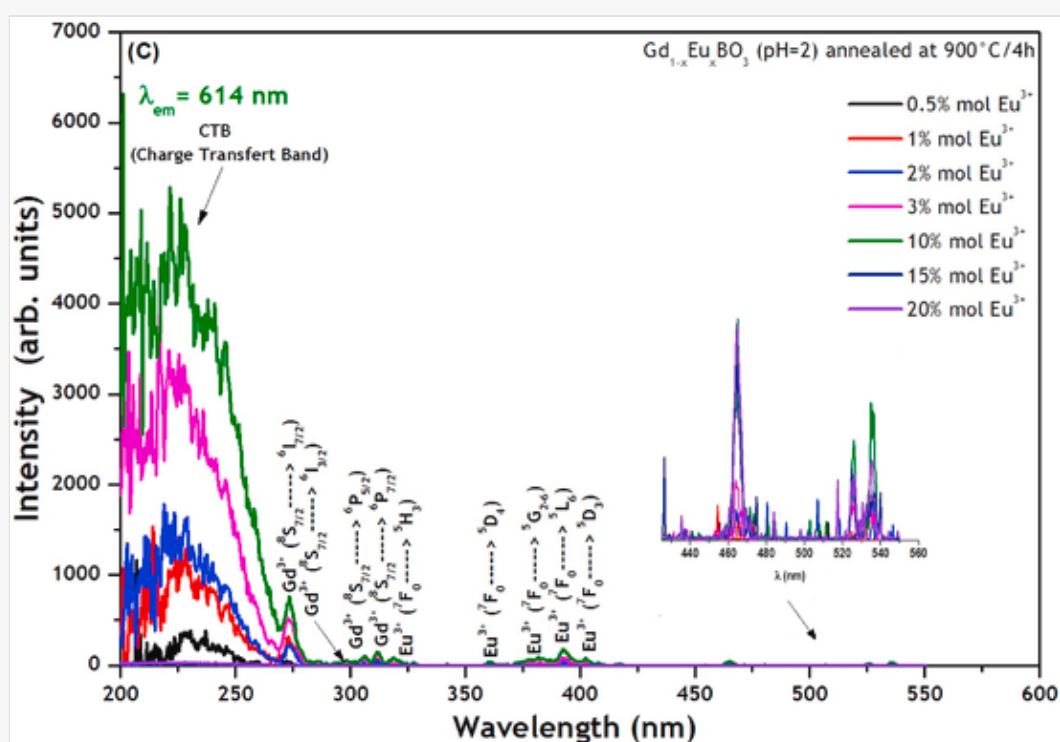
Fig. 7(b)



The room temperature excitation spectrum (Fig. 7(c)) of the $\text{GdBO}_3:10\% \text{Eu}^{3+}$, annealed at 900°C for 4 h was recorded by monitoring the emission peak at 614 nm. Fig. 7 (c) reveals a broad and strong excitation band in the 200–265 nm range, which indicates the charge-transfer (CT) transition between O^{2-} and Eu^{3+} . The weak excitation peak observed at 275 nm is assigned to the $^8\text{S}_{7/2} \rightarrow ^6\text{I}_J$ transitions of Gd^{3+} ion. The presence of this peak indicates that it exist an energy transfer from Gd^{3+} to Eu^{3+} [44]. Several peaks from 300 to 540 nm (Fig. 7 (c) inset) are likely to have originated from the intra-4f forbidden transitions of Eu^{3+} ions. the excited Gd^{3+} states than the $^5\text{D}_0$ state of Eu^{3+} ; consequently, Gd^{3+} ions can be play an intermediate for the energy transfer and sensitize Eu^{3+} luminescence.

alt-text: Fig. 7(c)

Fig. 7(c)

Excitation spectra of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ (pH = 2) powders, annealed at $900^\circ\text{C}/4\text{h}$

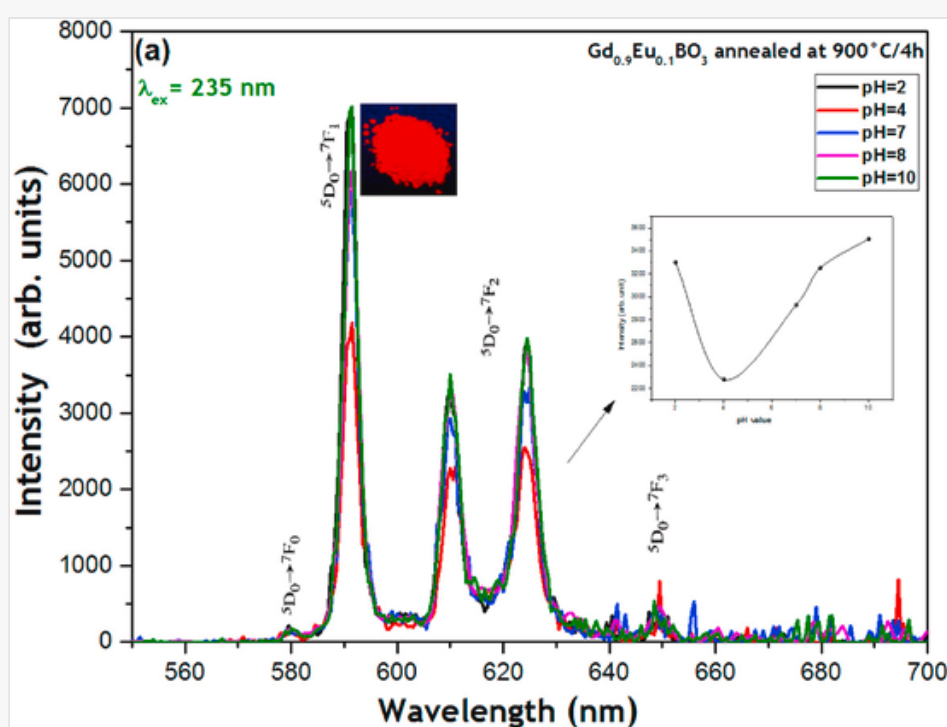
In fact, the charge-transfer (CT) transition process corresponds to the transfer of an electron from the 2p orbitals of the oxygen ligands to the empty 4f orbitals of the Eu^{3+} ion, giving rise usually to the charge transfer band (CTB). In most cases, CTB exhibits a stronger intensity than that of the 4f-4f interconfigurational transitions that appearing as narrow peaks with low intensities in excitation spectra. It is known that this interconfigurational transition of type ($\text{np}^q, 4\text{f}^6 \rightarrow \text{np}^{q-1}4\text{f}^7$) has a high absorption cross-section compared to that of 4f⁶-4f⁶ intraconfigurational transitions, which explain difference in intensity between them on the excitation spectra. Also, this process can induce a significant luminescence of the dopant ion Eu^{3+} doped oxides material if the population of radiative levels $^5\text{D}_J$ by a process of energy transfer is itself important [45].

The room temperature steady photoluminescence excitation and emission, of $\text{GdBO}_3:10\% \text{Eu}^{3+}$, for different pH value of solution namely: pH 2, 4, 7, 8 and 10 annealed at 900°C for 4 h duration were measured and displayed in Fig. 8(a and c). The emission spectra were measured under 235 nm excitation wavelength (Charge transfer Band (CTB)). The spectra show the characteristic features of Eu^{3+} emission, with orange $^5\text{D}_0 \rightarrow ^7\text{F}_1$

transition domination. From Fig. 8(a), one can see that intensity of the observed orange-red $^5D_0 \rightarrow ^7D_J$ transition emissions are influenced by the pH values of the solution. Indeed, with increasing the pH value from 2 to 4, the intensity decreases and increases from 4 to reach the higher intensity at pH = 10 (Fig. 8(a) inset), despite the presence of the highest H₂O absorption band in this sample (Fig. 6(c)), which quench usually the luminescence. The highest emission at pH = 10 compared to the other pH values can be explained by two ways: (i) the sample prepared at pH = 10 presents crystallite size among the smallest of the pH values (see Fig. 3(b)), which can leads the segregation of a large Eu³⁺ concentration ions on the surface of the crystallite, giving rise to the highest emission under UV (235 nm) excitation constituting the quenching concentration. (ii) Also, one can related this behavior to the structural purity, where the Gd_{0.9}Eu_{0.1}BO₃ sample prepared at pH = 10 presents a vaterite quasi-pure structure.

alt-text: Fig. 8(a)

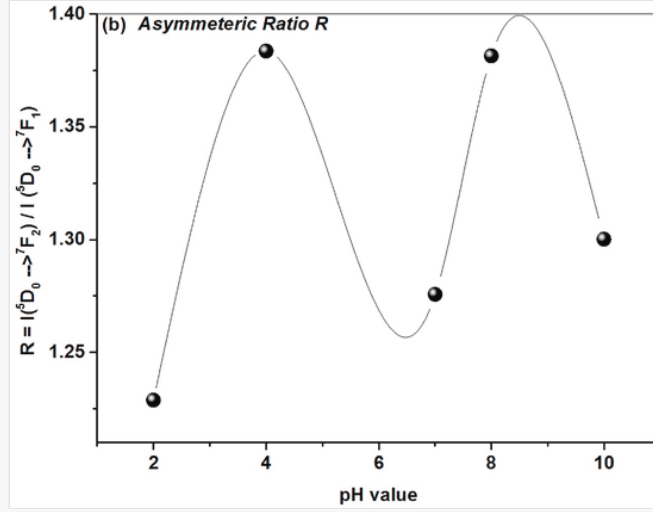
Fig. 8(a)



Emission spectra of Gd_{0.9}Eu_{0.1}BO₃ powders with different pH value of precursors suspension, annealed at 900 °C/4h.

alt-text: Fig. 8(b)

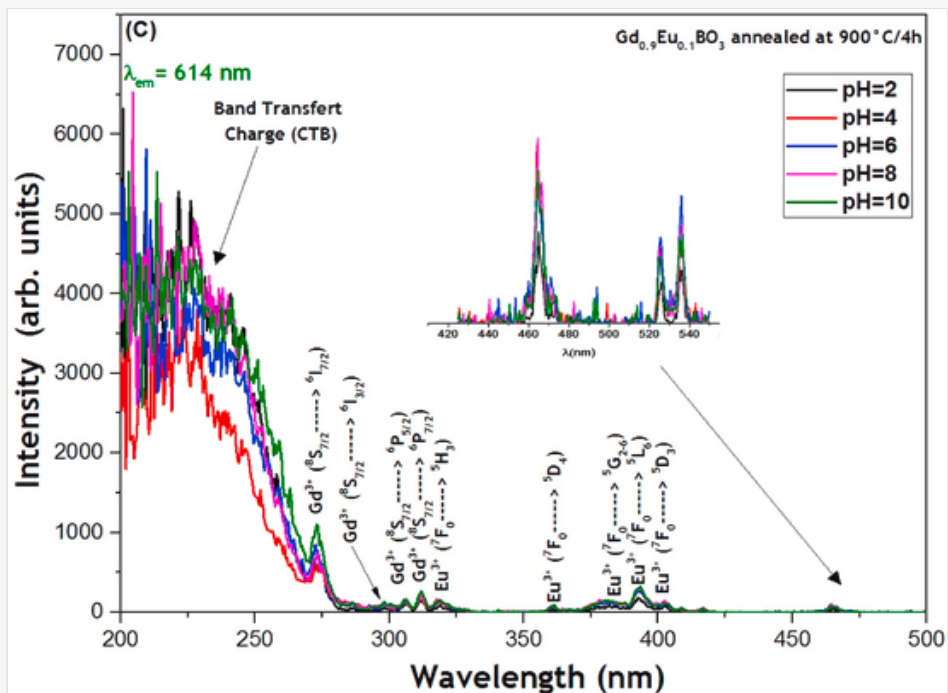
Fig. 8(b)



Asymmetry ratios for Eu^{3+} in GdBO_3 with various pH value of precursors suspension.

alt-text: Fig. 8(c)

Fig. 8(c)



Excitation spectra of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ powders with different pH value of precursors suspension, annealed at $900\text{ }^\circ\text{C}/4\text{h}$.

The variation of the asymmetry ratio parameter R in function of pH value of solution is displayed in Fig. 8(b). The values of R do not follow a well-defined trend and presents like zigzag shape between pH values. We can think that the value of R is very sensitive to several parameters, which can be changed by switching from one pH to another following, and thus influencing the symmetry of the site occupied by the Eu³⁺ ion in GdBO₃ host material. Also, from this figure, we can note that for some pH values we can promote the ⁵D₀→⁷F₂ red emission and can achieve better chromaticity, despite their lower luminescence compared to that pH = 10. Then, there is a tradeoff between the chromaticity and intensity of luminescence.

The excitation spectra for different pH value, annealed at 900 °C for 4 h were measured for the red emission (614 nm), and presented in Fig. 8(c). From this figure, the charge-transfer (CT) transition between O²⁻ and Eu³⁺ were observed between 200 and 265 nm in addition to Gd³⁺ excitation transition and those of intra-4f forbidden transitions (Fig. 8(c) inset) of Eu³⁺ ions.

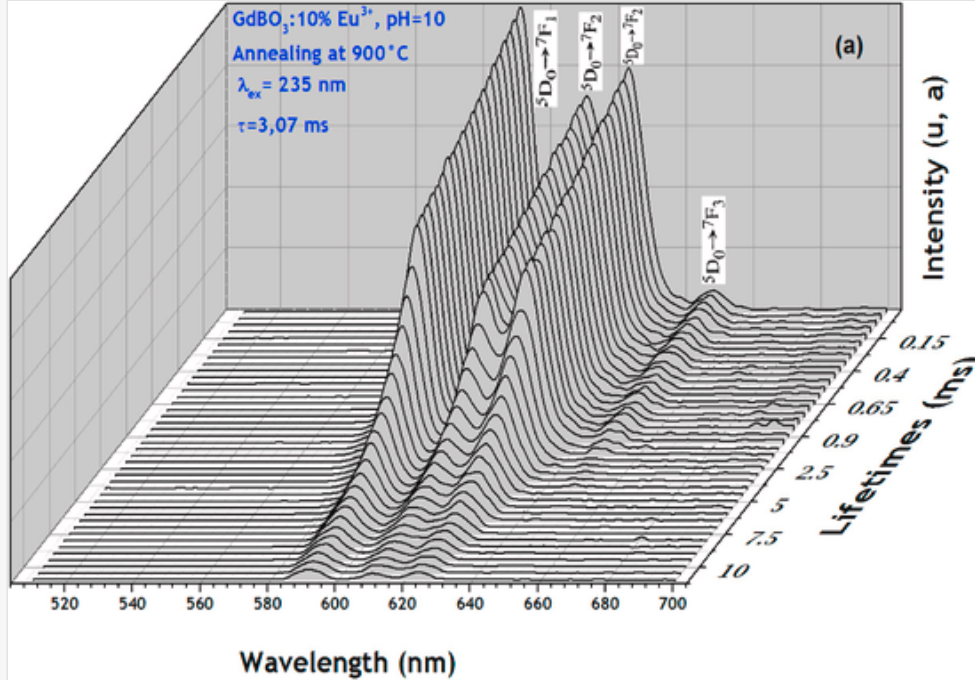
Fig. 9(a) displays the room temperature time-resolved phosphorescence emission spectra of GdBO₃:10 at% Eu³⁺, for pH = 10 annealed at 900 °C for 4 h duration as an example measured in phosphorescence mode. The extracted integrated photoluminescence decay curves of the ⁵D₀→⁷F₁ (591 nm) transition under UV excitation (CTB-235 nm) for different Eu³⁺ content and different pH value of solution is presented in Fig. 9(b) and Fig. 9(c) respectively. Referring to the decay curves, the effective decay times (τ_{AV}), which is considered as the experimental lifetime (t_{exp}) for ⁵D₀ level of Eu³⁺ can be determined using the following expression [33]:

$$\tau_{AV} = \frac{\int_0^{+\infty} tI(t)dt}{\int_0^{+\infty} I(t)dt} \quad (2)$$

Where $I(t)$ is the luminescence intensity at the considered time. The ⁵D₀ level lifetime as a function of Eu³⁺ concentration was presented in Fig. 9(b) inset. From Fig. 9(b) inset, one can observe that there are three regions with respect to the Eu³⁺ content. The first region situated between 0.5 and 2 at % Eu³⁺, in which the ⁵D₀ level lifetime decreases. The second one from 2 to 10 at % Eu³⁺, the ⁵D₀ level lifetime increases. The third last one, from 10 to 20 at % Eu³⁺ the lifetime decreases again. Several mechanisms could be used to explain the decreasing of ⁵D₀ luminescence lifetimes in these three regions. For the first and the third region, the reducing of the lifetime can be connected with cross-relaxation between the Eu³⁺ ions [46] and the existence of the parasitic phases, playing a quencher role. Indeed, the Eu³⁺ ions can interact in their excited states, and also, this interaction can happen between Eu³⁺ its excited state and Eu³⁺ neighboring ions in a ground state, which result in energy transfer and depopulation of Eu³⁺ ions in the ⁵D₀ state. For the third region, the increasing of the lifetime can be related to fact that GdBO₃ becomes the more quasi-purest hexagonal vaterite structure and presents the biggest crystallite size with respect to Eu³⁺ concentration.

alt-text: Fig. 9(a)

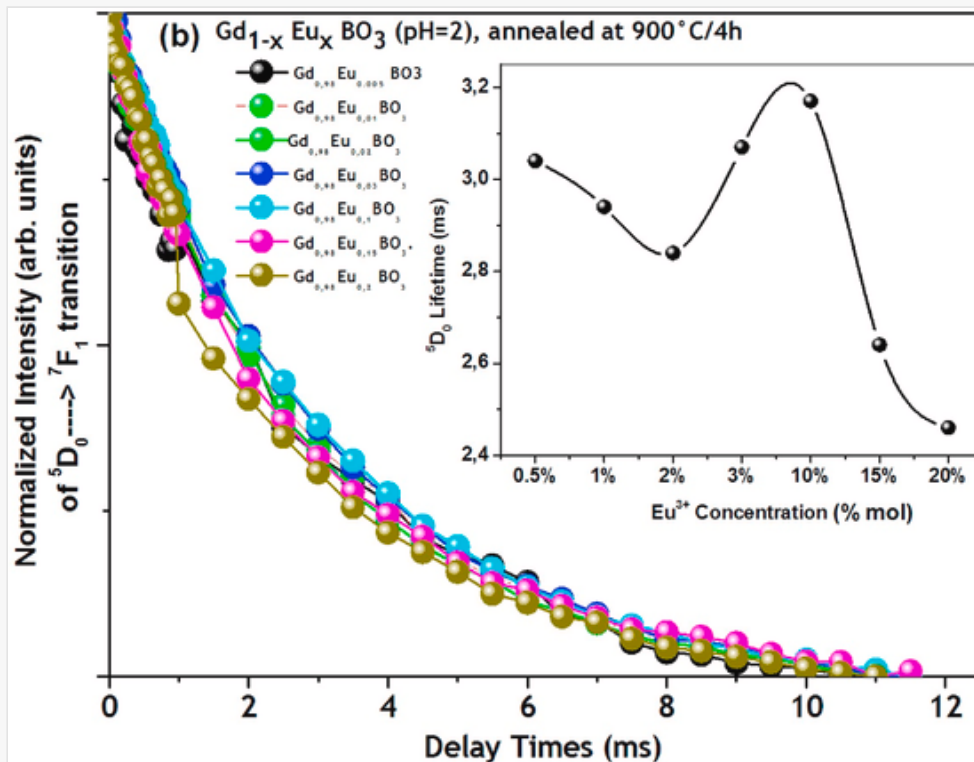
Fig. 9(a)



Time resolved emission spectra of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ (pH = 10) powder, annealed at $900^\circ\text{C}/4\text{h}$

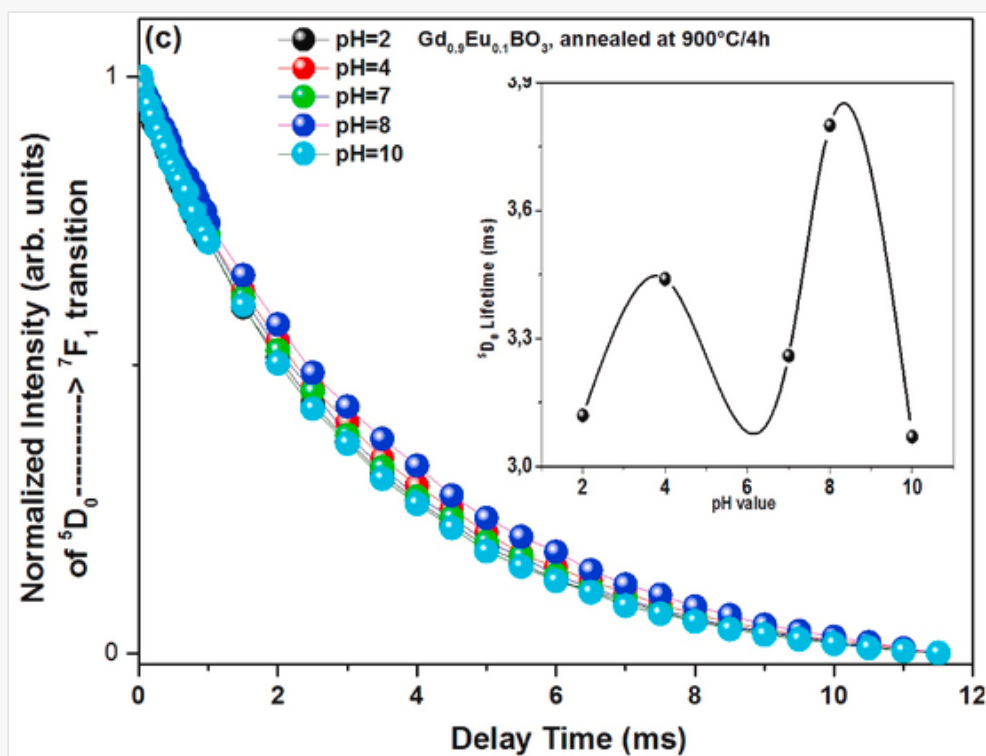
alt-text: Fig. 9(b)

Fig. 9(b)



Delay time curves of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition of $\text{Gd}_{1-x}\text{Eu}_x\text{BO}_3$ (pH = 2) powders, annealed at $900^\circ\text{C}/4\text{h}$.

Fig. 9(c)



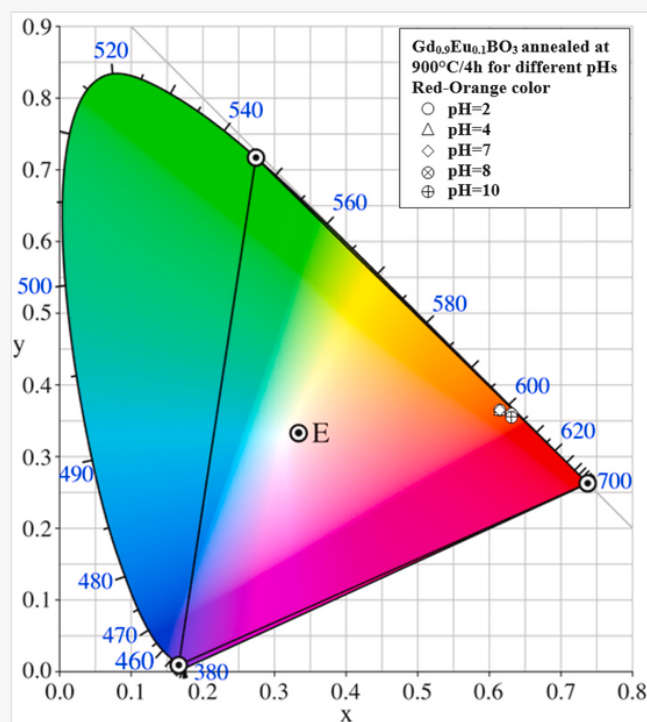
Delay time curves of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transition of $\text{Gd}_{0.9}\text{Eu}_{0.1}\text{BO}_3$ powders with different pH value of precursor suspension, annealed at $900^\circ\text{C}/4\text{h}$.

From Fig. 9(c) inset, one can see variation that ${}^5\text{D}_0$ level lifetime depends strongly on the pH value of solution. In fact, the value of ${}^5\text{D}_0$ level lifetime is connected to several related parameters such as the synthesized method, the phase purity, structure defects, atomic or molecular impurity in the host material and the crystallite size. Although $\text{GdBO}_3:\text{Eu}^{3+}$ samples prepared at pH = 2 and 10 present the purest hexagonal vaterite structure their ${}^5\text{D}_0$ level lifetime values are smallest compared to the intermediate pH values. One can think that the crystallite size is responsible for this smallest lifetime. Indeed, for pH = 10 the sample presents the highest orange-red emission and a small crystallite size. Another possible cause is the presence of the H_2O absorbed band with high quantity at pH = 10, which quenches the emission leading to a short lifetime.

Furthermore, in nano-size optical materials, several parameters such as the structure, the size have an important impact on the radiative and nonradiative lifetimes, energy transfer, and thermalization phenomena. In fact, it was found that rare earth doped nano-sized inorganic materials can exhibit prolonged lifetimes [47], which further explain the increases of the lifetime in the third region in Fig. 9(b) inset and the variation with respect of pH values.

In addition, for evaluating the performance of such $\text{GdBO}_3:\text{Eu}^{3+}$ nanophosphor, the chromaticity coordinate is one of the important parameters. Fig. 10 displays the color coordinates for $\text{GdBO}_3:10\%$ at $\% \text{Eu}^{3+}$, for different pH values of solution. The CIE coordinates of the synthesized samples are given in Table 2. It can be clearly seen from the figure and the table that most of the chromaticity coordinates are located in the orange-red region. In addition, one can see that CIE coordinates are closer to the NTSC (National Television Systems Committee) value of red phosphors (0.670, 0.330) [48]. These results indicate that this kind of phosphor can be useful in the development of LEDs applications.

Fig. 10

Chromaticity coordinates diagram at different pH values of precursors suspension of $Gd_{0.9}Eu_{0.1}BO_3$ annealed at $900\text{ }^\circ\text{C}/4\text{h}$

alt-text: Table 2

Table 2

i The table layout displayed in this section is not how it will appear in the final version. The representation below is solely purposed for providing corrections to the table. To preview the actual presentation of the table, please view the Proof.

(x,y)Chromaticity coordinates at different pH values of precursor suspension of $Gd_{0.9}Eu_{0.1}BO_3$ annealed at $900\text{ }^\circ\text{C}/4\text{h}$. The standard NTSC red color coordinates are presented in ref [48].

pH value	CIE Coordinate	
	x	y
2	0.6318	0.3678
4	0.6334	0.3662
7	0.6336	0.3660
8	0.6361	0.3634
10	0.6355	0.3641
NTSC	0.670	0.330

4 Conclusion

Intense orange-red emitting GdBO₃ orthoborate based nanophosphors doped with Eu³⁺ ion have been successfully synthesized via simple and low-cost, using acetylacetonate assisted aqueous sol-gel process. Systematic studies have been performed for different doping Eu³⁺ concentration and the pH value precursors suspension. It has been found that the Eu³⁺ ion content and pH value of solution have crucial influences on the structural, morphology and spectroscopic properties. The purest of the hexagonal GdBO₃ vaterite structure is governed by Eu³⁺ concentration and pH values. GdBO₃ sample prepared with 10% mol Eu³⁺ at pH = 2 present the purest *P6₃/mmc* hexagonal vaterite structure. Furthermore, the existence of (BO₃)³⁻ groups together with (B₃O₉)⁹⁻ shown by vibrational analysis, is discussed and related to the GdBO₃ calcite type or other vaterite structure namely triclinic and monoclinic. The crystallite size and morphology of the samples can be readily tailored by adjusting the pH and Eu³⁺ doping. Although, the crystallite size of GdBO₃:10% mol Eu³⁺ at pH = 10 is small, the corresponding particle size is largest. Under UV excitation (Charge Transfer Band (CTB)), all Eu³⁺-doped GdBO₃ samples show orange-red emission lines corresponding to transitions originate only from the excited ⁵D₀ levels to the ⁷F_J levels (J = 0–3) of Eu³⁺ ion. The variation of the emission intensities as well as ⁵D₀ lifetime in function of Eu³⁺ content and pH value are closely connected to morphology and crystallinity of GdBO₃ host nanomaterial. Besides, it has been found that the Eu³⁺ and pH values have an important influence on the asymmetry ratio parameter R between the red ⁵D₀→⁷F₂ and orange ⁵D₀→⁷F₁ transitions intensities. The chromaticity coordinates are located in the orange-red region with high luminescence efficiency and therefore, GdBO₃:Eu³⁺ nanophosphors make them potentially useful for optical devices.


CRediT authorship contribution statement

M. Seraiche: Investigation, Formal analysis, Data curation, Visualization, Writing - original draft, Writing - review & editing. **L. Guerbous:** Validation, Visualization, Project administration, Writing - original draft, Writing - review & editing. **R. Mahiou:** Methodology, Formal analysis, Writing - review & editing. **A. Potdevin:** Methodology, Formal analysis, Writing - review & editing.

Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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 The corrections made in this section will be reviewed and approved by a journal production editor. The newly added/removed references and its citations will be reordered and rearranged by the production team.

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Highlights

- GdBO₃:Eu³⁺ nanophosphors were prepared using acetylacetonate assisted aqueous sol-gel method; with two essential parameters: Eu³⁺ concentration and pH value.
- Effects of Eu³⁺ content and the pH value on the morphology, structural and luminescence spectroscopy properties have been investigated and highlighted.
- GdBO₃:qEu³⁺ hexagonal pure phase shows a correlation on the europium content and pH value.
- High orange-red color emission from ⁵D₀ levels to the ⁷F_J levels (J = 0–3) of Eu³⁺ ion was observed.
- It was found that 10% concentration of Eu³⁺ presents the quenching of emission and sample pH = 10 exhibit the highest emission intensity in GdBO₃.

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Answer: GdBO₃:Eu³⁺ nanophosphors were prepared using acetylacetonate assisted aqueous sol-gel route with Eu³⁺ content and pH value; · Effects of Eu³⁺ content and the pH value on the morphology, structural and luminescence properties were investigated. · GdBO₃:Eu³⁺ hexagonal pure phase shows a correlation on the europium content and pH value. · High orange-red color emission from ⁵D₀ levels to the ⁷F_J levels (J = 0–3) of Eu³⁺ ion was observed. · For 10 % Eu³⁺ content the quenching was found for pH = 10 exhibiting the highest emission intensity.